The Physico-chemical Constants of Binary Systems in Concentrated Solutions

VOLUME 3

SYSTEMS WITH METALLIC COMPOUNDS

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Preface to Volume 3

In this third volume, we have collected all the numerical data published about concentrated binary mixtures with at least one metallic compound.

To prepare this volume we have received substantial grant from the Union Minière du Haut-Katanga, which is especially interested in this kind of binary systems, and we wish to thank heartily this big Society for its support.

As in the former volumes, there are in some cases considerable discrepancies between the data given by different authors; this is due not only to some impurity of the chemical samples but also to the use of experimental methods or apparatus which are not always adequate. For example, even if the thermometers are accurately calibrated, some authors made the mistake of not taking into account the correction for the emergent column, which is positive or, for low temperatures, negative; with electrical thermometers, where the accuracy can be much greater, there are errors connected with the temperature coefficient of the resistant coils, if the thermostat is not very good; in viscosity measurements, the method of Poiseuille is accurate only if the diameter of the capillary, relative to its length, is adequate; and so on.

Today, all physical chemists who want to make accurate measurements should test their physical methods with great care.

Brussels
October, 1959

Jean Timmermans

Contents

N. TWO METALLIC SALTS

XL. TWO METALLIC SALTS WITH THE SAME ANION (Fluorides and Chlorides)

Lithium fluoride (LiF) + Sodium fluoride (NaF)

Bergman and Dergunov, 1941

%	f.t.
0	844
39	652 E
100	990

Lithium fluoride (LiF) + Potassium fluoride (KF)

Bergman and Dergunov, 1941

%	f.t.
0	844
-	492 E
100	856

Lithium fluoride (LiF) + Beryllium fluoride (BeF_2)

Thilo and Lehmann, 1949

mol%	f.t.	E	mol%	f.t.	E
0	845	-	47.5	400	350
10	780	465	48.7	390	355
20	700	470	50	365	-
27.5	570	465	52	-	360
30	500	472	53	-	365
31.5	-	462	57	-	367
33.3	475	-	60	-	365
36.0	475	(350)	63	-	360
38.0	453	(345)	67	417	370
40	445	340	70	417	365
42.5	440	340	73	420	365
54.0	420	340	80	431	370

Roy, Roy and Osborn, 1950 (fig.)

mol%	f.t.	E	mol%	f.t.	E
0	870	-	51	356	356
10	-	456	70	480	356
20	456	456	80	520	356
33	358	-	100	540	-
40	430	456			

(2 + 1)

See also page 2 .

Lithium fluoride (LiF) + Magnesium fluoride (MgF_2)

Tacchini, 1924

wt%	mol%	f.t.	m.t.	min.
0	0	840	-	-
9.42	8.05	830	818	12
19.62	16.98	810	785	25
29.35	25.83	785	750	35
39.30	35.20	750	720	30
44.25	39.90	725	700	25
49.21	44.54	700	685	15
54.15	49.69	685	-	-
59.1	54.79	675	-	-
63.65	59.55	669	-	-
68.2	64.27	670	-	-
73.70	69.99	682	-	-
78.79	75.71	690	-	-
88.5	86.59	1060	-	-
100	100	1400	-	-

Bergman and Dergunov, 1941

mol %	f.t.
0	844
33	742 E
100	1080

Haven, 1950 (fig.)

mol %	f.t.	mol %	f.t.
0	842	30	736
10	815	32	724 E
20	765		

mixed crystals .

mol %	f.t.	mol %	f.t.
4.6	700	0.3	500
2.7	650	0.04	400
1.4	600		

Counts, Roy and Osborn, 1953 (fig.)

mol%	f.t.	m.t.	solid sol
0	870	-	-
10	830	800	670
20	790	770	720
36	735	735	735
50	980	830	735
60	1080	910	730
80	1210	1090	715
100	1280	-	-

Novoselova, Simanov and Yarembash, 1952

%	f.t.	tr.t.I	(min.)	tr.t.II	(min.)	tr.t.III	tr.t.IV
0	845	-	-	-	-	-	-
9.0	831	464	1.5	-	-	-	-
13.5	820	462	4.0	-	-	-	-
13.6	-	458	-	-	-	-	-
18.6	812	461	5.0	-	-	-	-
23.8	773	460	6.0	-	-	-	-
25.0	774	457	6.0	-	-	-	-
25.6	784	459	7.0	-	-	-	-
30.4	722	465	9.0	-	-	-	-
31.7	782	460	10.0	-	-	-	-
36.0	682	465	-	-	-	-	-
37.8	662	465	-	-	-	-	-
40.1	603	461	-	-	-	-	-
41.8	571	457	-	-	-	-	-
43.1	570	458	-	-	-	-	-
46.0	485	461	12.0	-	-	-	-
47.1	562	460	-	-	-	-	-
47.5	560	460	-	-	-	-	-
47.5	-	460	12.0	-	-	-	-
49.4	-	460	8.5	-	-	330	-
51.7	478	460	7.5	306	1	-	-
52.4	450	-	-	348	3	327	-
52.9	452	-	-	348	4	-	214
53.8	448	-	-	347	4	-	-
54.7	433	-	-	340	3.5	330	216
55.8	432	-	-	359	3	319	-
57.8	428	-	-	361	4.5	316	-
58.9	422	-	-	355	-	-	-
61.3	390	-	-	351	4.5	-	-
62.6	390	-	-	341	6.0	-	-
63.8	379	-	-	357	7.0	322	-

%	f.t.	m.t.	E f.t	min.	E m.t.	min.	tr.t.	min.	tr.t.I	min.	tr.t II	min.
66.3	-	360	370	8.0	346	7.0	-	-	-	-	-	-
67.6	-	353	370	10.0	-	-	-	-	338	-	-	-
68.9	-	341	367	9.5	-	-	-	-	330	0.3	-	-
70.6	-	352	369	9.5	-	-	-	-	-	-	-	-
71.2	-	364	372	8.5	-	-	-	-	337	0.7	-	-
73.3	392	357	367	8.5	-	-	244	-	338	0.8	-	-
73.9	420	342	363	8.0	-	-	281	0.2	-	-	210	-
75.5	460	367	368	6.5	349	8.5	-	-	-	-	215	-
77.1	-	357	371	4.5	341	7.5	284	6.5	-	-	215	-
77.7	465	443	368	5.0	362	8.0	-	-	-	-	213	0.3
78.6	486	431	369	5.0	347	7.0	252	4.5	339	1.0	217	-
79.4	483	413	374	4.5	340	7.0	260	5.8	330	1.0	-	-
80.3	490	441	373	4.5	370	7.5	-	-	-	-	216	0.3
80.9	493	424	373	4.5	344	7.5	278	5.5	337	1.0	218	-
82.3	505	437	367	4.5	350	7	-	-	-	-	223	-
84.4	508	423	368	4.0	343	7.5	-	-	342	2.5	212	-
86.4	518	426	370	4.0	353	5.5	282	3.0	341	1.5	-	-
88.5	527	-	370	3.0	-	-	270	1.5	347	1.0	-	-
90.5	537	450	368	2.5	348	4.5	235	0.5	337	4.0	221	0.3
92.4	532	461	370	2.5	344	6	-	-	314	3.0	-	-
94.4	540	451	371	1.5	347	3.5	-	-	333	4.5	221	0.3
96.0	543	451	367	0.5	337	3.5	-	-	326	5.5	--	-
97.9	554	459	359	-	339	3	-	-	-	-	218	0.6
99.9	567	446	353	-	-	-	-	-	387	-	217	0.7
100	577	-	-	-	-	-	-	-	-	-	-	-

Lithium fluoride (LiF) + Calcium fluoride (CaF_2)

Bergman and Bichkova, 1955

mol%	f.t.	mol%	f.t.
0	848	24.3	816
1.6	840	26.6	842
3.1	833	29.1	867
4.8	825	31.6	895
6.4	817	34.3	916
8.2	810	37	940
9.9	804	39.9	962
11.8	796	42.9	986
13.7	791	46	1012
15.7	782	49.3	1036
17.7	776	52.7	1062
19.8	769 E	56.3	1090
22	792		

Lithium fluoride (LiF) + Barium fluoride (BaF_2)

Bichkova and Bergman, 1956

mol%	f.t.	mol%	f.t.
0	848	50.0	812tr.t.
5.0	838	55.0	824
10.0	825	60.0	826
12.5	816	62.5	826
15.0	810	64.0	827 (1+1)
20.0	800	65.0	837
25.0	783	70.0	884
30.0	768	75.0	916
32.5	760	75.0	920 tr.t.
32.0	760 E	77.5	964
35.0	766	80.0	964
40.0	780	85.0	1020
45.0	795	90.0	1088

Lithium fluoride (LiF) + Aluminum fluoride (AlF_3)

Pushin and Baskov, 1913

mol%	f.t.	E	min.	
0	870	-	-	
5	828	710	-	
8	790	703	3.0	(3+1)
12.5	730	700	4.6	
14.5	706	706	11.4	
19.0	750	700	2.2	
22.5	783	690	0.6	
27	780	661	1.2	
30.5	737	689	4.3	
32.5	728	689	7.4	
37	-	691	8.7	

Fedotiev and Timofeev, 1932 (fig.)

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	860	860	27.5	782	710
2.5	848	715	30.0	768	710
5.0	830	715	35.0	722	"
10.0	784	715	36.0	710	710E
12.5	755	715	37.5	726	710
15.0	715	715 E	40.0	755	"
17.5	750	715	45.0	853	"
20.0	770	715	50.0	-	"
25.0	790	715		(3+1)	

Sodium fluoride (NaF) + Potassium fluoride (KF)

Kurnakov and Zhemchuzhni, 1907

mol%	f.t.	E	mol%	f.t.	E
0	997.0	-	55.4	699	699
5.84	967	700	60	716	699
11.33	948	-	65	768	-
22.44	899	-	73.8	771	697
33.16	847	-	80	788	-
42.66	798	-	85.0	812	-
48.2	770	699	89.67	812	-
53.1	743	-	100	837	-

Kordes, 1926

mol %	f.t.
0	997
60	700
100	837

Bergman and Nagornii, 1943

%	f.t.	m.t.	%	f.t.	m.t.
100	854	852	55	744	743
90	824	823	50	774	772
80	791	790	45	801	799
70	756	756	40	825	824
65	735	734	30	878	877
60	717	716	20	919	918

Bergman and Dergounov, 1945 (fig.)

mol %	f.t.	mol %	f.t.
100	850	40	830
80	780	20	950
60	710 E	0	1000
50	770		

Sholokhovich, 1955

E : 60 mol % 708°

Sodium fluoride (NaF) + Beryllium fluoride (BeF_2)

Novoselova, Levina and al., 1944

wt%	mol%	f.t.	E	min.
0	0	1024	-	-
2.15	1.92	1006	580	1
5.51	4.95	935	590	1
11.30	10.21	827	580	1.5
17.79	16.20	715	580	1.0
20.00	18.25	664	590	1.7
21.50	19.66	624	580	2.0
23.11	21.16	620	580	2.3
24.98	22.93	580	580	2.0
29.19	24.91	604	580	-
32.00	29.60	609	580	1.7
35.32	32.79	620	-	-
37.90	35.27	536	353	1.0
39.89	37.22	520	360	1.7
41.08	38.39	517	360	-
42.58	39.86	470	360	2.1
43.21	40.46	456	353	--
45.80	43.02	377	360	-
53.00	50.19	440	360	-
56.81	54.03	470	-	-
60.41	57.69	528	-	-
64.21	61.57	-	-	-
69.53	67.09	602	-	-
75.75	73.63	682	-	-
85.82	84.37	734	-	-
94.97	94.47	744	-	-
100.00	100.00	780	-	-

(i+2)

wt%	tr.t. (2+1) I	tr.t. (2+1) II	min. I	min. II
2.15	-	220	-	0.7
5.51	-	220	-	0.7
11.30	-	-	-	-
17.79	322	220	0.7	1.0
20.00	-	220	-	1.0
21.50	322	220	0.7	1.6
23.11	-	220	-	1.3
24.98	322	220	0.7	1.6
29.19	340	-	-	-
32.00	320	220	1.0	2.0
35.32	330	222	2.0	2.3
37.90	330	222	1.5	2.0
39.89	330	220	1.5	2.0
41.08	330	-	-	-
42.58	330	220	1.3	0.7
43.21	323	-	1.0	-

(2+1)

wt%	tr.t. (1+1) I	tr.t. (1+1) II	min.	tr.t. BeF_2 I	tr.t. BeF_2 II
45.80	-	-	-	-	-
53.00	378	-	-	-	430
56.81	360	-	-	-	-
60.41	370	-	-	-	-
64.21	360	290	8	-	-
69.53	360	295	10	-	411
75.75	374	284	5	-	-
85.82	390	284	5	-	459
94.97	372	-	-	-	473
100.00	-	-	-	528	425

Thilo and Schröder, 1951

mol%	f.t.	E	mol%	f.t.	E
0	990	-	38	480	-
10	920	555	39	415	-
15	875	565	40	375	-
18	837	560	40	348	-
21	785	560	42	342	-
25	700	560	44.3	340	-
26	682	560	47	358	335
28	618	570	50	372	-
29	595	565	52.5	360	-
31	560	560	55	355	-
32	570	560	61.4	340	-
33.	578	-	75	345	-
35	540	-	83	345	-

(1+1) (2+1) (3+2)

Roy, Roy and Osborn, 1950 (fig.)

mol%	tr.t.	mol%	tr.t.
50	381	80	510
56	375 E	90	530
70	470	100	540

Sodium fluoride (NaF) + Magnesium fluoride (MgF_2)

Bergman and Dergunov, 1941 and 1945 (fig.)

mol%	f.t.	mol%	f.t.
0	1000	50	1030
10	960	64	987
25	816	80	1100

(1+1)

Sodium fluoride (NaF) + Barium fluoride (BaF_2)

Grube, 1927

%	f.t.	E	min.
0	992	-	-
5	984	-	-
10	977	817	47
15	970	813	70
20	961	823	93
25	952	823	117
30	942	825	135
35	932	823	162
40	919	816	188
45	907	821	210
50	893	821	234
55	872	825	259
60	856	825	290
65	838	822	337
70	823	823	425
75	847	823	365
80	889	819	297
85	964	814	220
90	1050	816	145
95	1155	-	-
100	1287	-	-

Sodium fluoride (NaF) + Calcium fluoride (CaF_2)

Fedotiev and Iljinski, 1923

mol%	wt%	f.t.	E
100	100	1360	-
70	81.3	1090	810
50	65.0	950	"
32.5	47.2	810	"
20	31.7	880	"
10	17.1	935	"
0	0	990	-

Grube and Henne, 1930

%	f.t.	E	%	f.t	E
100	1330	-	40	840	803
80	1123	796	30	896	804
70	1010	798	20	937	801
60	902	801	10	961	798
50	812	803	0	992	-
47.2	-	803			

Sodium fluoride (NaF) + Zinc fluoride (ZnF_2)

Puschin and Baskov, 1913

mol%	f.t.	E	min.
0	1040	-	-
5.5	983	660	2.1
13.5	940	665	4.3
23.0	905	682	5.8
100	872	-	-

Sodium fluoride (NaF) + Cadmium fluoride (CdF_2)

Puschin and Baskov, 1913

mol%	f.t	E	min.
0	1040	-	-
9	995	645	3.1
13.5	970	650	4.6
23	895	655	5.5
34.5	770	655	7.8
45	678	660	9.3
54.5	746	660	5.6
65	885	655	2.4
90	1040	640	1.0
100	1110	-	-

Sodium fluoride (NaF) + Lead fluoride (PbF_2)

Puschin and Baskov, 1913

mol%	f.t.	E	min.
0	1040	-	-
10	990	540	1.7
20	930	542	2.3
30	872	545	3.4
40	825	538	-
50	740	540	-
60	628	538	-
66.5	560	540	-
69.5	575	540	5.0
76.5	648	540	3.5
85	710	531	1.9
100	855	-	-

Gladushchenko and Bergman, 1955

mol%	f.t.
0	990
63.1	498 E
100	860

Sodium fluoride (NaF) + Aluminum fluoride (AlF_3)

Puschin and Baskov, 1913

mol%	f.t.	E	min.	tr.t.	min.
0	1040	1040	-	-	-
2.5	1030	930	0.6	-	-
7	1000	935	2.6	600	0.7
11	978	932	3.8	-	-
14.5	948	935	7.5	-	-
18.5	973	920	4.9	603	3.3
19	990	923	4.4	600	3.9
21	1005	910	-	602	-
23.5	1017	-	-	600	-
25	1023	-	-	600	4.1
28	1010	742	0.5	-	-
30	995	740	5.8	603	1.0
30.5	985	740	4.1	600	-
31.5	950	743	-	-	-
32	945	740	7.4	-	-
35	860	740	1.12	-	-
40	-	737	-	- (3+1)	-

Fedotieff and Iljinski, 1913

mol%	f.t.
100	1000
14	885 E
0	990

Lorenz, Jobs and Eitel, 1913

%	f.t.	E	min.
100	999	-	-
37.6	992	884	2.7
32.4	972	884	5.0
26.8	926	887	7.9
25.0	917	889	8.1
21 7	904	895	10.8
0	985	-	-

Fedotieff and Iljinski, 1913 and 1924

mol%	wt%	f.t.	E_1	min.	E_2	min.	tr.t	min
0	0	990	-	-	-	-	-	-
3	5.82	980	-	-	-	-	565	-
4	7.69	970	875	0.2	-	-	"	0.2
5	9.52	965	875	2	-	-	"	0.8
10	18.18	920	880	2.8	-	-	"	1
14	24.56	885	835	8	-	-	"	1.2
15	26.10	900	885	6.7	-	-	560	1.3
20	33.33	955	830	4.1	-	-	565	2.1
23	37.40	990	875	0.3	-	-	"	2.5
25	40.00	1000	-	-	-	-	"	2.5
30	46.17	980	725	5.5	-	-	"	1
35	51.85	915	"	6.5	-	-	-	0.2
37.4	54.33	855	"	7.5	685	0.5	-	-
40.4	57.48	730	"	5	675	1.5	-	-
42.3	59.45	720	-	-	675	2.5	-	-
45	62.12	700	-	-	680	6	-	-
46.6	63.50	685	-	-	685	10.5	-	-
47.4	64.32	695	-	-	"	9	-	-
48.5	65.32	770	-	-	"	-	-	-
50	66.67	-	685	-	-	8	-	-

Elchardus, 1938

%	σ
t = f.t. +20°	
0	12.8
26	14.5
27.5	13.4
29	10.6

Sodium fluoride (NaF) + Ferric fluoride (FeF_3)

Puschin and Baskov, 1913

mol%	f.t.	E	mol%	f.t	E
0	1040	-	25	975	895
10	1025	885	30	945	890
15	1005	890	35	-	892

Sodium fluoride (NaF) + Praseodymium fluoride (PrF_3)

Dergunov, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	992	30	780 E
10	950	35	800
15	920	45	880
20	835	50	945

Potassium fluoride (KF) + Rubidium fluoride (RbF)

Dergunov and Bergman, 1948

mol%	f.t.	mol%	f.t.
100	780	45	802
95	780	40	808
90	781	35	814
85	782	30	820
80	782	25	827
75	783	20	832
70	784	15	838
65	785	10	845
60	788	5	850
55	792	0	856
50	797		

Mixed crystals

Potassium fluoride (KF) + Magnesium fluoride (MgF_2)

Remy and Seeman, 1940

mol%	f.t.	min.	tr.t.I	min.
100	1260	3.0	-	-
90	1215	-	1020	0.5
80	1145	-	1020	1.25
70	1045	-	1019	5.0
68	1020	7.0	-	-
60	1068	-	1010	1.25
50	1080	3.0	-	-
45	1072	-	1060	-
40	1055	-	1040	-
35	1004	-	940	-
33.3	970	-	869	0.5
30	890	-	850	-
25	850	-	840	0.25
21.5	840	-	782	1.5
20	822	-	782	2.0
17.5	805	-	782	2.75
15	792	-	781	3.0
12.5	780	3.25	-	-
10	793	-	780	1.5
8.5	797	-	780	1.25
5	810	-	780	0.5
0	848	2.0	-	-

(1+1)

mol%	tr.t.II	min.	tr.t.III	min.	tr.t.IV	min.
60	840	0.1	-	-	-	-
50	840	0.25	-	-	-	-
45	863	0.5	840	0.5	800	1.75
40	860	0.75	800	-	-	-
35	862	1.5	840	0.5	800	2.0
33.3	842	-	812	0.75	800	-
30	840	0.25	780	0.5	-	-
25	830	-	780	0.75	-	-

Bergman and Dergunov, 1945 (fig.)

mol%	f.t.	mol%	f.t.
0	850	68	1010 E
10	786 E	80	1150
22	870 tr.t	100	1270
50	1060 (1+1)		

Dergunov and Bergman, 1948

mol%	f.t.		mol%	f.t.	
0	856	-	35	1040	(1+2)
5	837	-	40	1070	"
10	806	-	45	1086	"
12.5	785	E	50	1090	(1+1)
15	812	(2+1)	55	1085	"
17.5	836	"	60	1076	"
20	856	"	65	1052	"
22.5	872	tr.t.	68	1028	E
24	886	(1+2)	70	1045	-
25	910	"	75	1094	-
30	983	"	80	1145	-

De Vries and Roy, 1953

mol%	t			E	
	I	II	III		
0	850	-	-	-	
10.0	803	-	-	776	
20.0	-	-	850	778	
25.0	903	-	852	782	
33.3	998	-	846	780	
40.0	1036	-	844	778	
43.0	1043	973	-	-	
46.0	1039	1007	842	777	
48.0	1058	1008	843	778	(1+1)
50.0	1070	-	-	-	
52.0	1068	1006	-	-	
55.0	1048	1007	-	-	
63.5	1114	1012	-	-	
75.0	-	1008	846	778	

Potassium fluoride (KF) + Beryllium fluoride (BeF_2)

Borzenkova, Novoselova and al., 1956

%	f.t.	E	tr.t.		
			1	2	3
0	846	-	-	-	-
5.2	810	700	-	-	-
11.0	760	700	-	-	-
14.0	720	700	-	-	-
16.8	735	700	-	-	-
21.3	-	715	745	-	-
23.0	-	-	740	-	-
25.0	-	-	740	-	-
28.5	795	-	735	-	-
28.8	800	-	-	685	-
32.0	770	400	-	700	390
35.4	740	407	-	695	400
41.3	505	410	-	-	395
43.1	450	400	-	-	395
43.5	408	405	-	-	400
44.7	415	-	-	-	400
46.9	402	316	-	-	-
48.3	399	399	-	-	-
50.7	368	338	-	241	-
52.0	-	349	-	282	-
53.1	-	338	-	274	157
54.1	-	346	-	270	168
56.9	-	334	-	-	-
59.4	-	334	-	-	155
61.0	-	342	-	274	164
62.6	-	338	-	293	164
64.8	-	340	-	278	185
68.1	-	330	-	-	-
69.8	470	346	-	283	178
73.7	450	346	-	-	-
75.8	490	-	-	274	179
82.1	539	347	-	-	197
87.7	533	-	-	-	-
93.4	544	-	-	-	-
100	545	-	-	-	-

* 1 = equilibrium (3+1)-(2+1)
2 = transition of (2+1) I-II;2b = decomposition of (1+2)
3 = transition Of (1+1) I-II; 3b = probable transition of (1+2)

X-ray analysis of mixed crystals

Potassium fluoride (KF) + Barium fluoride (BaF_2)

Puschin and Baskov, 1913

mol%	f.t.	E	min.
0	885	-	-
7	870	735	-
13	865	735	2.6
23.5	840	740	3.0
35	807	745	5.0
50	775	750	10.6
60	780	750	8.0
70	885	750	5.5
80	1005	750	2.4
90	1105	735	1.2
100	1280	-	-

Potassium fluoride (KF) + Lead fluoride (PbF_2)

Gladushchenko and Bergman, 1956

mol%	f.t.	mol%	f.t.
0	850	36	546
8.2	816	39	486
11.2	803	40.4	460 E
15	778	42	504
17	766	45	530
21.2	736	54.7	584
26	682	58.8	610
40.2	632	100	826

Potassium fluoride (KF) + Nickel fluoride (NiF_2)

Wagner and Balz, 1952

mol%	f.t.	E	tr t
0.0	848	-	-
4.9	826	790	-
9.2	797	797	-
13.4	847	801	-
15.2	868	798	-
18.4	895	796	-
21.8	920	798	-
22.8	930	795	-
23.8	938	796	930
26.0	951	797	930
26.6	953	796	929
27.9	979	795	931
30.3	1016	790	927
32.3	1044	790	930
34.8	1053	791	931
37.9	1086	777	927
43.7	1111	-	927
46.7	1124	-	920
50.5	1130	-	-
53.0	1127	1077	-
58	1113	1084	-
63.5	1100	1083	-
71.5	1153	1086	-
75	1185	1084	-

Potassium fluoride (KF) + Aluminum fluoride (AlF_3)

Pushin and Baskov, 1913

mol%	f.t.	E	min.	tr.t.
0	885	-	-	-
4	856	838	1.7	-
5	850	837	2.8	-
10	885	837	3.8	-
13	939	845	2.8	302
18	987	835	-	302
20	998	835	0.5	302
23.5	1020	835	0.3	305
24.3	1027	-	-	300
24.5	1033	-	-	300
25.5	1028	500	0.7	-
28.5	1010	545	2.4	-
35	930	572	8.6	-
37	870	560	9.5	-
39	-	565	18.8	-
40	-	570	19.6	-
41	-	563	13.1	-
42.5	-	568	8.3	(3+1)

Fedotiev and Timofeev, 1932

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	873	873	20.0	995	840
0.5	872	860	25.0	1025	840
1.0	871	850	27.5	1016	565
1.3	870	840	30.0	1008	565
2.5	860	840	40.0	809	565
5.0	847	840	44.9	565	565
6.0	840	840 E	45.3	575	575tr
7.5	859	840	47.5	575	575t.
10.0	892	840	48.0	590	575
15.0	947	840	50	650	575
(1+1)	(3+1)		52	575	575

Potassium fluoride (KF) + Cerium fluoride (CeF_3)

Puschin and Baskov, 1913

mol%	f.t.	E	min.
0	885	-	-
3	860	640	0.3
6	838	643	0.6
9.5	809	643	1.0
19.5	708	648	3.4
21.5	692	650	3.8
24.5	-	660	6.1

Potassium fluoride (KF) + Samarium fluoride (SmF_3)

Dergunov, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	860	25	882
6	825	35	850
14	758 E	40	825 E
20	860	44	900
		50	980

(3+1)

Potassium fluoride (KF) + Praseodymium fluoride (PrF_3)

Dergunov, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	860	30	690
10	805	35	740
20	700	40	900
24	610 E	44	975

Potassium fluoride (KF) + Erbium fluoride (ErF_3)

Dergunov, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	860	25	1012
5	830	30	948
12	756 E	37	730 E
20	960	45	950

(3+1)

Rubidium fluoride (RbF) + Magnesium fluoride (MgF_2)

Dergunov and Bergman, 1948

mol%	f.t.	solid phase	mol%	f.t.	solid phase
0	780	-	35	830	(1+1)
5	764	-	37	862	"
10	743	-	40	884	"
15	714	-	45	904	"
19	686	E	50	912	"
20	716	(2+1)	55	905	"
22	748	"	60	893	"
25	770	"	62.5	883	E
27	780	"	65	920	-
30	788	"	70	990	-
32	790	"	75	1060	-
33.5	792	E	80	1126	-
34.5	810	(1+1)			

Remy and Seemann, 1940

mol%	f.t.	min.	E	min.
90	1192	-	875	2.0
80	1100	-	875	2.25
75	1060	-	875	3.25
70	980	-	875	3.75
65	875	4.25	-	-
63	875	4.0	-	-
60	878	-	875	3.5
55	880	-	875	1.0
50	881	3.5	-	-
45	860	-	780	0.5
40	840	-	780	1.0
37.5	815	-	780	1.5
35	783	2.25	-	-
33.3	784	2.25	-	-
30	782	-	700	0.5
25	753	-	700	0.75
24	738	-	700	1.0
22	700	1.5	-	-
20	722	-	700	1.5
15	740	-	700	1.25
10	754	-	700	0.75

E		f.t.
1	65 mol%	875
2	35 mol%	781
3	22 mol%	700
(1+1)		889
(2+1)		784

Rubidium fluoride (RbF) + Aluminum fluoride (AlF_3)

Pushin and Baskov, 1913

mol%	f.t	E	min.	tr.t.
0	833	-	-	-
3	813	790	0.5	-
4.5	805	792	1.6	320
7	798	795	3.8	-
11.5	845	790	2.5	354
18	900	790	0.51	353
23	964	790	0.2	348
26.5	967	437	-	323
29	930	505	3.6	322
31	900	515	5.9	-
35	735	530	10.5	-
36.5	602	527	12.5	-
39.5	-	542	16.7	-
39.7	-	560	15.6	-
40.9	-	560	13.8	-
41.5	-	560	10.2	-

(3+1)

Rubidium fluoride (RbF) + Erbium fluoride (ErF_3)
Dergunov, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	786	20	990
5	715	25	1034
7.5	732	30	990
10	760	35	870
15	920		

(3+1)

Rubidium fluoride (RbF) + Samarium fluoride (SmF_3)

Dergunov, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	786	25	916
6	750	30	880
12	700 E	36	800E
20	860	42	880

(3+1)

Rubidium fluoride (RbF) + Praseodymium fluoride (PrF_3)

Dergunov, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	786	30	750
6	760	37	660 E
15	640 E	45	825
20	740	50	910
25	797		

(3+1)

Cesium fluoride (CsF) + Aluminum fluoride (AlF_3)

Pushin and Baskov, 1913

mol%	f.t.	E	colour
0	715	-	-
2	702	670	-
5.5	-	685	green
9	718	672	"
15	785	675	" (3+1)
22.5	820	-	-
25	823	-	-
28.5	755	488	red
31.5	725	490	"

Cesium fluoride (CsF) + Praseodymium fluoride (PrF_3)

Dergunov , 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	685	32	855
7	654 E	38.5	783 E
12	775	44	770
18	870	50	960
25	920		

(3+1)

Cesium fluoride (CsF) + Samarium fluoride (SmF_3)

Dergunov, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	685	25	972
6.5	645	30	940
10	750	37	865E (3+1)
15	875	42	925

Cesium fluoride (CsF) + Erbium fluoride (ErF_3)

Dergunov , 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	685	20	995
6	648 E	25	1048
10	760	30	1010
15	890	35	900

(3+1)

Silver fluoride (AgF) + Zinc fluoride (ZnF_2)

De Vries and Roy, 1953

mol%	f.t.	m.t.	sat.t.
0	435	-	-
10.0	411	395	341
15.0	400	-	354
20.0	452	480	341
33.3	-	433	-
40.0	655	-	305
50.0	710	-	-
54.0	667	630	-
60.0	-	661	-
65.0	687	639	-
66.7	712	641	-
80.0	838	766	-

Beryllium fluoride (BeF_2) + Magnesium fluoride (MgF_2)

Counts, Roy and Osborn, 1953 (fig.)

mol%	f.t.	E	mol%	f.t.	E
0	540	-	40	1000	528
5	528	528	50	1050	528
20	850	528			

Beryllium fluoride (BeF_2) + Barium fluoride (BaF_2)

Kirkina, Novoselova and Simanov, 1956.

mol%	f.t.	m.t.	E h	E c	tr.t I.	
100	1280		-		-	
95	-	-	920	915	-	-
90	-	-	920	910	340	-
85	1160	1160	915	912	345	330
82	1100	1085	918	915	360	-
80	1075	1060	920	918	330	340
77	1000	1010	918	915	400	-
75	-	960	915	915	-	-
73	-	-	920	916	-	-
70	-	-	915	913	350	-
67	970	970	920	918	-	-
65	1010	1000	918	915	340	335
62	1050	1045	920	915	-	-
60	1065	1050	915	912	335	-
57	1075	1070	920	915	-	-
45	1080	1080	920	918	-	345
42	1080	1080	-	-	-	-
50	1080	1080	-	-	360	-

mol%	f.t.	m.t.	L_1 + h	L_2 c	E	tr.t. II
48	1080	1080	-	-	-	-
45	1075	1070	780	-	-	-
43	1050	1050	780	775	-	-
40	1030	1020	777	775	615	-
37	990	990	780	780	-	-
35	955	955	780	780	600	-
30	900	900	775	775	610	-
28	860	857	780	770	600	-
25	825	815	780	765	610	-
23	800	795	775	775	-	-
20	-	-	775	775	615	-
18	-	-	775	770	-	-
16	-	-	775	775	-	-
15	-	-	780	770	605	835
13	-	-	777	770	-	-
10	-	-	775	770	600	835
7	-	-	775	765	600	-
5	765	-	-	-	600	-

h = by heating
c = by cooling

Beryllium fluoride (BeF_2) + Calcium fluoride (CaF_2)

Counts, Roy and Osborn, 1953 (fig.)

mol%	f.t.	E	mol%	f.t.	E
89	-	890	30	860	-
70	1180	890	20	780	495
50	990	890	11	495	495
42	890	890	0	543	- (2+1)

Magnesium fluoride (MgF_2) + Calcium fluoride (CaF_2)

Beck, 1908

%	f.t.	E	min.
0	1264	-	-
10	1157	948	10.5
25	1105	945	11.3
40	1008	948	13.5
50	954	947	21.8
65	1086	944	14.5
80	1219	943	8.0
100	1330	-	-

Kordes, 1926

%	f.t.
0	1200
43.5	954 E
100	1325

Fuseya, Mori and Imamura, 1933 (fig.)

mol%	f.t.	mol%	f.t.
0	1248	60	1050
20	1150	80	1230
40	1000	100	1403
48	970		

Magnesium fluoride (MgF_2) + Barium fluoride (BaF_2)

Fuseya, Mori and Imamura, 1933 (fig.)

mol %	f.t.	mol %	f.t.
100	1353	60	980
90	1100	40	1100
80	920	20	1180
70-65	912 E	0	1248

Calcium fluoride (CaF_2) + Strontium fluoride (SrF_2)

Rumpf, 1930

Mixed crystals

%	c	%	c
0	5.449	75.5	5.712
25.5	5.539	81.6	5.726
38.4	5.576	91.65	5.759
50.65	5.624	100	5.781
66.5	5.675		

c = lattice constant in Å

Calcium fluoride (CaF_2) + Barium fluoride (BaF_2)

Fuseya, Mori and Imamura, 1933 (fig.)

mol%	f.t.	mol%	f.t.
100	1353	40	1280
80	1320	20	1350
60	1280	0	1403
50	1277		

Mixed crystals.

Bukhalova and Bergman, 1951

mol%	f.t.	mol%	f.t.
100	1280	45	1030
65	1100	40	1056
60	1060	35	1083
55	1032	0	1411
50	1022		

Calcium fluoride (CaF_2) + Aluminum fluoride (AlF_3)

Fedotiev and Iljinski, 1923

mol%	wt%	f.t.	m.t.
0	0	1360	-
20	29.8		820
30	31.6	1020	"
35	36.7	930	"
37.5	39.3	-	"
40	41.8	830	"
45	46.9	870	"
47.5	49.2	870	815
50	51.9	875	805
52	53.8	885	810
55	56.8	880	850
58	59.8	890	860

Calcium fluoride (CaF_2) + Yttrium fluoride (YF_3)

Vogt, 1914

%	f.t.	%	f.t.
0	1384	30	1389
10	1386	40	1367
20	1408	60	1269

Mixed crystals, with a maximum

Barium fluoride (BaF_2) + Strontium fluoride (SrF_2)

Beck, 1908

%	f.t.	%	f.t.
20	1146	70	1200
30	1188	100	1280
50	1175		

Strontium fluoride (SrF_2) + Lanthanum fluoride (LaF_3)

Ketelaar and Willems, 1937

mol%	d
0	4.18
10	4.26
20	4.41
30	4.46
50	4.47

Thorium fluoride (ThF_4) + Uranium fluoride (UF_4)

Dawson, 1951

t	χ	t	χ
5.3 mol%		11 mol%	
-183	2.903	-183	5.66
-74	1.437	-73	2.822
+27	0.936	+30	1.837
57.5	0.814	82	1.579
25 mol%		50 mol%	
-183	10.26	-183	16.92
-74	5.94	-73	10.11
-27	4.24	25	7.38
55	3.76	61	6.62
75 mol%		90 mol%	
-183	20.60	-183	22.27
-74	12.60	-73	13.41
+22	9.22	+28	10.01
55	8.47	61	9.21
100 mol%			
-183	22.39		
-76	14.61		
+27	10.67		
59	9.92		

Thorium fluoride (ThF_4) + Plutonium fluoride (PuF_4)

Dawson, 1952

t	χ	t	χ
100%		77.8%	
-183	9.482	-183	9.168
-72	7.384	-73	6.829
+28	5.589	+27	4.878
56	5.125	56.5	4.543
97	4.677	105.5	4.049
142.5	4.293	139.5	3.718
180	3.984	168	3.494
52.0%		34.0%	
-183	7.249	-183	5.745
-72	4.949	-73	3.874
+28	3.512	+27	2.527
100	2.907	61	2.306
169.5	2.449	89	2.036
		145	1.783
		169.5	1.489
3.6%		2.0%	
27	0.142	27	0

Sodium fluoberyllate (Na_2BeF_4) + Rubidium fluoberyllate (Rb_2BeF_4)

Toropov and Grebenshchikov, 1956.

mol%	f.t.	m.t.	tr.t. I -II
100	807	-	692
95	795	780	685
90	770	739	674
80	718	-	667
70	665	-	-
60	594	-	-

tr.t. II-III	II + III + L	E	tr.t. III -?
-	-	-	-
533	-	-	-
562	-	-	-
-	584	-	-
-	581	-	430
-	585	482	312

I,II and III are polymorfic forms of Rb_2BeF_4

Sodium aluminum fluoride (Na_3AlF_6) + Potassium aluminum fluoride (K_3AlF_6)

Naray-Szabo and Sigmond, 1941

%	f.t.	%	f.t.
0	1001	47.5	941
10	986	49	941
15	974	50	941
20	964	55	943
25	953	56.74	942
30	946	60	941
32.5	941	65	951
33.5	941	67.5	953
35	942	70	950
36	939	71	951
37	941	73	948
38	939	75	951
39	936	77	951
40	938	80	945
45	943	85	958
46	943	90	968
		100	991

Lithium chloride (LiCl) + Sodium chloride (NaCl)

Kangro and Wieking, 1938

mol%	p	p_1	p_2
		860°	
74.4	3.0 ± 0.4	1.1	1.9
50.7	3.2 ± 0.5	2.2	1.0
25.5	4.4 ± 0.4	3.7	0.7
		950°	
74.3	9.6 ± 0.6	3.6	1.0
50.8	10.7 ± 0.6	7.5	3.2
25.6	14.6 ± 0.6	11.6	3.0
		990°	
74.3	16.6 ± 0.7	6.1	10.5
50.7	18.6 ± 0.8	12.4	6.2
25.7	23.9 ± 1.0	21.0	2.9

Zhemchuzhni and Rambach, 1910

mol %	f.t.	m.t.	tr.t.
0	614	-	-
2.36	608	602	-
8.83	590	580	-
16.22	570	562	-
22.56	557	554	294
27.00	552	552	300
29.16	554	552	305
34.75	564	556	314
40.96	585	567	312
48.93	612	586	303
54.32	635	605	300
60.60	660	617	-
66.80	686	640	-
74.40	716	665	-
81:34	743	695	-
89.72	776	735	-
96.02	803.5	782	-
100	819	-	-

Gromakov and Gromakova, 1953

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	605	605	45	593	555
10	580	565	50	615	-
20	560	549	60	655	585
25	549	547	70	698	617
(27.5)	546	min.	80	733	670
30	549	546	(90)	769	-
35	560	547	100	800	800
40	575	549			

Akopov, 1956.

%	f.t.	%	f.t.
0	606	36	572
5	600	38	580
10	588	40	588
15	578	42	596
18	567	45	608
20	564	50	630
21.5	553 E	55	654
22	562	60	676
25	558	65	697
28	556	70	718
30	558	75	738
32	564	100	800
34	567		

tr.t. 28.5% 565°
33% 575°

Lithium chloride (LiCl) + Potassium chloride (KCl)

Zhemchuzhni and Rambach, 1910

mol%	f.t.	E	mol%	f.t.	E
0	614	-	44.70	390	352
1.85	605.5	350	51.40	448	352
6.85	582	350	56.10	492	352
12.81	552	351	61.66	539	352
19.30	512	352	68.08	594	352
24.80	477	352	76.28	653	351
30.60	435	352	86.17	713	350
35.54	395	352	95.00	764	349
40.47	352	352	100	790	-

Richards and Meldrum, 1917

mol%	f.t.	mol%	f.t.
100	773	42.3	363
65.7	600	41.5	361
60.5	552	41.3	362
57.4	527	40.6	372
52.2	477	38.7	388
44.2	385	35.6	412
42.7	367	0	601

Keitel, 1925

mol%	f.t.	E	min.
0	607	-	-
15	542	359	150
18.5	520	359	200
36.68	395	"	330
44.71	387	"	370
57.57	525	"	290
70.17	623	"	205
77.69	667	"	160
88.00	718	"	110
100	776	-	-

Elchardus and Lafitte, 1932

mol%	f.t. cooling	m.t. cooling	f.t. heating	m.t. heating
0	607	-	503	-
13.41	542	354	544	-
25.37	468	352	467	352
28.82	444	353	443	354
38.22	378	354	375	355
44.20	396	355	-	352
50.38	460	354	-	354
55.70	520	354	-	355

E: 41.7 359°

Dombrovskaya, 1933 (fig.)

mol%	f.t.	mol%	f.t.
0	605	60	485
20	480	80	620
42	358 E	100	775

Gromakov and Gromakova, 1953

mol%	f.t.	E.	
0	605	605	
10	564	456	
20	513	356	
30	449	356	
35	414	356	(1+1)
40	370	356	
(41.5)	356	-	
45	400	356	
50	452	356	
55	500	356	
60	540	356	
70	615	356	
(90)	733	-	
100	775	775	

Bergman, Kislova and Posypaiko, 1954

%	f.t.	%	f.t.
0	606	55	516
10	560	60	564
20	510	65	604
30	448	70	638
40	368	75	664
45	462	80	690
50	472	100	774

E : 42% 348° tr.t.: 565°

Karpachev, Stromberg and Podchainova, 1935

t	d	t	d	t	d
		80 mol %			
677	1.570	770	1.523	852	1.474
698	.558	791	.508	873	.468
729	.546	816	.493	883	.460
750	.535				
		60 mol %			
648	1.580	711	1.540	795	1.490
668	.568	735	.530	820	.478
691	.554	757	.519	846	.466
		50 mol %			
430	1.650	573	1.606	669	1.563
514	.635	592	.598	680	.555
539	.623	627	.582	704	.546
550	.618				
		40 mol %			
426	1.670	525	1.618	603	1.580
447	.656	552	.605	628	.570
483	.639	585	.592	661	.556
		20 mol %			
582	1.571	648	1.517	744	1.475
568	.567	673	.504	772	.465
611	.538	716	.487		

Van Artsdalen and Yaffe, 1955

t	d	t	d	t	d
		0 %			
620.3	1.4975	655.9	1.4821	737.8	1.4468
631.6	.4924	676.9	.4730	752.5	.4403
638.2	.4899	685.9	.4689	754.1	.4396
638.6	.4821	713.6	.4570	780.5	.4282
		18.23 %			
532.6	1.5751	592.0	1.5458	647.5	1.5187
561.7	.5605	605.8	.5390	660.5	.5123
577.2	.5529	626.3	.5291	660.6	.5122
589.7	.5466				
		29.64 %			
455.9	1.6245	516.4	1.5942	560.5	1.5715
456.7	.6242	533.6	.5850	577.0	.5633
471.0	.6171	534.3	.5847	602.1	.5504
495.2	.6046				
		41.20 %			
395.0	1.6766	481.7	1.6313	538.0	1.6012
395.3	.6764	482.7	.6308	555.2	.5922
440.0	.6528	504.5	.6188	579.8	.5791
445.0	.6505	518.2	.6119	593.1	.5722
477.7	.6331				
		59.55 %			
585.0	1.5958	645.4	1.5610	686.5	1.5378
588.3	.5932	648.5	.5592	746.0	.5054
613.4	.5794				
		80.04 %			
690.0	1.5614	795.1	1.4998	844.9	1.4715
719.2	.5439	795.8	.4994	845.5	.4711
761.2	.5196	834.6	.4773	853.3	.4670
		100 %			
779.5	1.5219	827.4	1.4943	907.2	1.4479
781.8	.5211	870.4	.4688	939.0	.4292
809.4	.5058				

Karpachev, 1935

t	η 100 mol%	90 mol%	70 mol%	60 mol%
550	-	-	-	2020
600	-	-	1800	1660
650	-	-	1520	1390
700	-	1280	1260	1140
750	-	1060	1100	990
800	1080	870	950	890
850	960	780	870	790
900	920		-	-

t	50 mol%	25 mol%	0 mol%	
400	-	5500	-	-
450	-	4210	-	-
500	3760	3080	-	-
550	2790	2850	7060	-
600	2050	1840	1660	-
650	1600	1400	1360	1600
700	1190	1130	1120	1210
750	1030	930	940	990
800	880	780	820	810
850	-	690	670	720

Benrath and Tesche, 1920

t	ϰ 0%	10%	20%	30%	40%
290	0.146	0.122	0.090	0.062	0.047
295	"	.142	.110	.080	.060
300	0.191	.167	.139	.102	.085
305	"	.196	.173	.131	.101
310	0.240	.234	.202	.173	.135
315	"	.282	.260	.226	.184
320	0.322	.384	.341	.312	.240
325	"	.539	.441	.413	.342
330	0.414	.962	.713	.601	.491
335	"	1.610	-	.912	.778
340	0.535	2.520	-	1.511	1.262
345	"	5.030	-	3.710	3.570
350	0.775	56.490	-	11.070	11.050

t	50%	60%	70%	80%	90%
290	0.022	0.012	0.006	0.005	-
295	.034	.024	.011	.016	-
300	.059	.036	.022	.026	-
305	.081	.054	.034	.037	0.011
310	.104	.078	.051	.053	.021
315	.140	.103	.080	.096	.039
320	.194	.199	.125	.161	.078
325	.271	.295	.160	.286	.136
330	.398	.345	.324	.547	.246
335	.631	.557	.531	1.820	-
340	1.130	1.094	.916	19.63	8.50
345	2.890	2.440	1.840	-	-
350	18.68	12.58	12.97	-	-

t	100%	t	100%
580	0.005	670	0.0137
590	.010	680	.0185
600	.015	690	.0233
610	.025	700	.0288
620	.035	710	.0378
630	.045	720	.0492
640	.060	730	.0667
650	.081	740	.0865
660	.106	780	21.0000 ?

Karpachev, 1935

100%		90%		30%	
t	ϰ (in mhos)	t	ϰ	t	ϰ
780	2.13	730	2.19	630	2.07
795	2.14	740	2.14	710	2.15
800	2.17	760	2.18	730	2.23
810	2.17	765	2.19	740	2.27
845	2.25	782	2.23	765	2.35
863	2.29	790	2.24	795	2.43
875	2.33	810	2.30	825	2.58
895	2.37	820	2.34	865	2.62
930	2.40	830	2.36	-	-
940	2.41	860	2.42	-	-
80%		**70%**		**50%**	
690	2.15	610	2.01	600	2.23
705	2.20	620	2.07	620	2.30
730	2.29	630	2.09	635	2.33
745	2.33	635	2.15	645	2.45
760	2.39	660	2.23	660	2.55
790	2.49	690	2.29	690	2.70
840	2.56	700	2.35	715	2.76
860	2.59	715	2.39	750	2.96
880	2.62	725	2.43	770	2.98
900	2.67	735	2.47	800	3.01
920	2.70	747	2.49	-	-
930	2.71	-	-	-	-
40%		**30%**		**25%**	
660	2.71	500	2.46	650	3.46
690	2.83	520	2.56	660	3.55
720	2.96	535	2.64	690	3.68
750	3.10	550	2.72	740	3.86
790	3.20	570	2.85	770	3.96
810	3.35	580	2.94	780	3.99
840	3.54	608	3.09	790	4.07
-	-	628	3.18	-	-
-	-	640	3.26	-	-
-	-	650	3.32	-	-
-	-	675	3.40	-	-
-	-	685	3.47	-	-
-	-	705	3.56	-	-
20%		**10%**		**0%**	
500	2.98	600	4.28	630	5.63
540	3.19	611	4.36	650	5.77
565	3.31	636	4.43	660	5.77
590	3.48	666	4.60	730	6.19
630	3.66	692	4.75	780	6.30
660	3.72	704	4.78	820	6.54
680	3.81	734	4.91	850	6.78
695	3.88	774	5.03	870	6.83
730	3.99	789	5.19	-	-
740	4.05	824	5.37	-	-
760	4.19	-	-	-	-

Van Artsdalen and Yaffe, 1955

t	κ (mho cm^{-1})	t	κ (mho cm^{-1})
	0%		
622.8	5.6923	712.5	6.2253
634.4	5.8962	741.0	.3704
643.9	5.8922	760.3	.4626
667.8	6.0083	783.3	.5399
691.1	6.1185		
	18.23 mol%		
521.1	3.0443	612.2	3.5708
529.4	.1075	622.5	.6268
529.7	.1087	626.4	.6438
551.8	.2440	642.5	.7214
562.6	.3054	670.8	.8492
579.0	.4006	671.2	.8500
599.7	.5121		
	29.64 mol%		
450.1	1.9074	544.1	2.5765
450.4	.9123	547.5	.5954
450.7	.9239	563.9	.6881
465.5	2.0773	585.1	.7969
	41.20 mol%		
389.8	1.1819	486.0	1.7964
394.0	.2040	514.8	.9524
441.5	.5162	522.2	.9987
443.0	.5314	535.2	2.0655
476.1	.7290	553.5	.3375
	59.55 mol%		
589.0	1.8480	662.6	2.2028
634.2	2.0726	685.8	.2864
641.6	.1144	701.2	.3602
657.5	.1783	737.5	.4734
	80.04 mol%		
706.9	2.0765	765.4	2.2644
707.2	.0773	804.7	.3735
726.5	.1331	807.6	.3842
760.0	.2529	832.3	.4397
765.0	.2634	855.7	.4975
	100 mol%		
778.9	2.1628	850.9	2.3609
782.4	.1816	851.2	.3618
787.0	.1984	865.0	.3901
790.0	.2074	877.5	.4174
799.0	.2342	889.0	.4476
800.5	.2407	892.5	.4484
804.1	.2478	905.5	.4745
823.0	.3020	906.7	.4777
846.0	.3507	925.1	.5156

Lithium chloride (LiCl) + Rubidium chloride (RbCl)

Zhemchuzhni and Rambach, 1910

mol%	f.t.	E	mol%	f.t.	E
0	614	-	51.59	368	312
2.19	604.5	311	59.10	453	"
7.74	579.5	312	66.77	527	"
15.96	535	"	74.56	586	"
23.79	484	"	82.10	635	311
28.10	450	"	89.05	672	311
33.07	415	"	95.92	706.5	310
38.85	356	"	100	726	-
44.75	312	"			

Richards and Meldrum, 1917

mol%	f.t.	mol%	f.t.
100	709	43.4	323
65.6	534	42.0	320
47.2	339	41.2	327
46.4	327	30.6	455
44.8	324	0	601
(1+1)			

Keitel, 1925

mol%	f.t.	E	min.
100	722	-	-
84.6	653	316	30
63.84	520	320	90
57.63	452	321	130
50	362	321	160
45	319	315	125
40	339	316	220
30	425	315	180
20	504	312	120
0	607	-	-

Lithium chloride (LiCl) + Cesium chloride (CsCl)

Korreng, 1915

%	f.t.	%	f.t.
100	635	79.87	351
97.28	603	78.22	348
92.25	520	76.45	341
90.25	474	74.58	336
88.81	447	72.57	343
85.62	385	66.49	395
84.30	376	49.80	512
82.91	368	0	609
81.43	359		

%	I	II	I -(1+2)		II -(1+2)	
	tr.t.	min.	tr.t.	min.	tr.t.	min.
100	479		-	-	-	-
97.28	475	115	372	85		
92.25	475	60	373	150	346	40
90.25	-	-	380	175	356	50
88.81	-	-	380	190	356	60
85.62	-	-	381	70	360	60
84.30	-	-	-	-	362	50
82.91	-	-	-	-	-	-

%			(1+1)	
	t	min.	t	min.
88.81	333	40	-	-
85.62	340	90	-	-
84.30	342	110	-	-
82.91	345	120	-	-
81.43	347	200	-	-
78.22	-	-	329	70
76.45	-	-	332	150
74.58	-	-	330	240
72.57	-	-	332	570
66.49	-	-	332	440
49.80	-	-	333	360

Richards and Meldrum, 1917

mol%	f.t.	mol%	f.t.
100	645	48.3	354
74.5	529	46.2	345
63.7	442	43.3	334
54.3	368	38.9	339
52.0	357	32.3	407
49.9	356	0	601

(1+1)

Dergunov, 1951

mol%	f.t.	mol%	f.t.
100	640	47.5	358
95	626	45	346
90	606	42.5	335
85	582	41.5	332 E
80	566	40	346
75	542	37.5	368
70	508	35	390
67.5	488	32.5	412
65	465	30	435
62.5	442	25	477
60	417	20	518
58	397 tr.t.	15	554
57.5	396	10	578
55	387	5	595
50	367	0	606

Lithium chloride (LiCl) + Thallium chloride (TlCl)

Sandonnini and Aureggi, 1911

mol%	f.t.	E	min.
0	602	-	-
5	583	343	25
10	557	339	30
20	520	343	50
30	480	342	90
40	440	342	100
50	397	342	130
60	350	342	130
70	373	345	90
90	410	340	30
95	420	340	-
100	429	-	-

Lithium chloride (LiCl) + Silver chloride (AgCl)

Sandonnini, 1911

mol%	f.t.	min.	m.t.	E
0	602	-	-	-
10	580	-	548	-
20	565	-	-	468
30	540	30	-	469
40	526	50	-	468
50	510	80	-	469
60	493	60	-	469
65	480	40	460	468
70	474	20	459	-
75	466	-	457	-
80	461	-	456	-
85	460	-	456	-
90	459	-	456	-
95	457	-	455	-
100	455	-	-	-

Salstrom, Kew and Powell, 1936

mol%	activity AgCl 500°	AgCl 600°	LiCl 500°	LiCl 600°
100.0	1.000	1.000	0.00	0.00
80.4	0.884	0.824	0.436	0.399
69.0	.791	.775	.614	.567
57.3	.741	.712	.687	.658
46.9	.690	.649	.740	.729
25.2	.529	.487	.851	.841
13.6	.392	.352	.912	.906
0	.000	.000	1.000	1.000

Lesnykh and Bergman, 1956

mol %	f.t.	mol %	f.t.
100	455	25	520
75	460	0	604
50	475		

Lithium chloride (LiCl)+ Cuprous chloride (CuCl)

Sandonnini, 1911

mol%	f.t.	m.t.	E	min.
0	602	-	-	-
10	573	560	-	-
20	546	510	-	-
30	512	-	424	20
40	472	-	425	40
50	450	-	425	60
55	440	-	424	80
60	428	412	423	20
65	420	409	-	-
70	415	409	-	-
75	413	409	-	-
80	403	-	-	-
85	409	-	-	-
90	414	-	-	-
95	417	-	-	-
100	422	-	-	-

Korreng, 1914

%	f.t.	m.t.	E	min.
100	425	-	-	-
97.68	418	397	-	-
95.46	412	383	-	-
93.34	405	380	-	-
89.66	404	384	-	-
87.51	405	381	-	-
85.09	409	386	-	-
82.37	415	398	-	-
79.26	427	-	417	440
75.69	431	-	415	430
66.69	480	-	415	350
53.87	550	-	413	220
34.17	581	-	413	100
10.95	606	584	-	-
0	609	-	-	-

Lithium chloride (LiCl) + Beryllium chloride ($BeCl_2$)

Schmidt, 1929

mol%	f.t.	E		min.
0	630	-		-
15	551	350		7
25	480	350		10
30	438	350		13
35	392	350	tr.t.	15
37.5	370	292	(2+1)	8
41	348	295		3
44	343	297		4
50	325	300		10
55	306	300		15
65	328	300		13
75	350	300		8
85	368	295		4
91	380	293		3
100	400	-		-

E : 56 mol% 300°

Lithium chloride (LiCl) + Magnesium chloride ($MgCl_2$)

Sandonnini, 1913

mol%	f.t.	E	mol%	f.t.	E
0.0	602	-	56.8	580	-
5.58	590	583	70.2	595	579
11.3	586	-	80.0	602	595
16.2	580	-	89.0	658	646
30.5	577	-	100.0	712	-
39.7	570	-			

Markov, Delimarskii and Panchenko, 1955

mol%	e	τ	mol%	e	τ
		718°			
100	2.531	-0.69	29.3	2.582	-0.56
66.6	.535	-0.58	22.7	.595	-0.516
55.3	.540	-0.564	14.8	.618	-0.526
44.7	.546	-0.58	10.4	.638	-0.55
36.6	.559	-0.51	4	.697	-0.48

e = electromotive force, in volts

τ = temperature coefficient . 10^3 (in volt/degree)

Lithium chloride (LiCl) + Calcium chloride ($CaCl_2$)

Sandonnini, 1913

mol%	f.t.	t(1+1)	mol%	f.t.	t(1+1)
0.0	602	-	36.4	492	445
4.3	590	-	41.6	537	440
9.2	575	-	47.0	620	
14.3	560	430	60.0	704	432
21.1	540	438	77.4	738	-
27.5	525	440	90.0	763	-
30.1	512	445	100.0	772	-
33.4	496	445			

Grube and Rüdel, 1924

%	f.t.	E	(1+1)
0	612	-	-
22.5	581	-	440
31.6	568	-	442
34.5	552	488	442
43.8	548	495	444
47.2	542	494	446
49.2	528	496	444
52.9	521	496	443
55.2	514	500	440
58.6	507	496	442
63.6	505	498	440
65.9	512	498	442
70.7	536	498	442
72.4	548	496	439
74.4	555	500	442
76.3	568	498	444
79.7	587	496	436
86.0	640	493	442
88.8	675	495	441
91.3	697	496	436
93.7	719	-	434
95.9	740	-	438
100	774	-	-

Golubeva and Bergman, 1954

mol%	f.t.
0	609
35.2	475 E
100	774

Gromakov and Gromakova, 1955

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	603	603	41	496	480
4.5	595	570	54	570	480
13.5	552	524	68.5	680	480
27.5	513	473	100	777	777
39	-	480			

Lithium chloride (LiCl) + Strontium chloride ($SrCl_2$)

Sandonnini, 1913

mol%	f.t.	E	min.
0.0	602	-	-
5.0	597	471	30
10.0	582	473	50
20.0	560	472	90
30.0	536	471	110
40.0	500	473	150
50.0	581	473	120
60.0	667	473	100
70.0	740	472	70
80.0	790	473	60
90.0	832	473	40
95.0	832	471	20
100.0	860	-	-

Golubeva and Bergman, 1955

E : 32.3 mol% 488°

Lithium chloride (LiCl) + Barium chloride ($BaCl_2$)

Sandonnini, 1913

mol%	f.t.	E	min.
0.0	602	-	-
5.0	592	507	30
10.0	580	508	40
20.0	560	511	80
30.0	532	510	150
40.0	556	510	120
50.0	649	510	90
60.0	713	512	70
70.0	775	510	50
80.0	825	510	40
90.0	876	511	30
95.0	903	502	
100.0	960	-	-

Gromakov and Gromakova, 1955

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	602	-	35	562	-
5	590	550	50	668	514
15	554	514	100	960	-
30	-	514			

Lithium chloride (LiCl) + Zinc chloride ($ZnCl_2$)

Evseiev and Bergman, 1951

mol%	f.t.		mol%	f.t.	
100	320	B	45	360	IIA
93.8	302	"	40	396	"
89.0	308	"	36	425	"
83.0	303	"	32	450	"
77.0	294	E	29	483	"
74	296	(2+1)	22	517	"
68	306	"	18	540	"
63	316	"	15	552	"
56	330	"	12	562	"
53	335	"	9.6	570	tr.t.
49	344	"	9	576	IA
45	352	E (2+1)	8	584	"

Lithium chloride (LiCl) + Cadmium chloride ($CdCl_2$)

Ferrari and Baroni, 1928

mol%	f.t.	mol%	f.t.	
0	603	64.0	502	
8.5	593	66.2	494	
15.0	578	69.8	500	
19.5	564	73.2	507	
25.8	555	75.2	516	
34.8	535	75.9	512	
43.0	514	77.3	501	
45.5	505	81.0	494	(2+3)
48.5	502	84.8	497	
52.0	506	86.2	513	
56.0	518	88.5	532	(3+1)
59.0	521	92.3	551	
60.2	522	100.0	560	
61.1	511			

Dergunov, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	562	40	520
80	535	20	560
60	502	0	600

Gromakov, 1950

mol%	f.t.	m.t.	mol%	f.t.	m.t.
		mixed crystals			
100	565	-	52.0	514	502
90.5	540	527	42.0	533	515
82.0	523	510	36.0	544	521
71.0	510	502	28.0	556	536
62.0	500	-	0	606	606

Lesnykh, Bergman and Bakun, 1956 (fig.)

mol%	f.t.	mol%	f.t.	
81.9	515	17.7	555	
60	500	14.3	561	tr.t.
42.9	515	11.2	570	
33.4	545	0	600	

Lithium chloride (LiCl) + Stannous chloride ($SnCl_2$)

Rack, 1914

mol%	f.t.	E	min.
100	239	-	-
98.84	257	211	80
97.58	223	213	160
96.21	-	215	400
94.71	240	216	380
91.27	327	211	340
87.04	375	214	320
71.91	491	216	240
52.80	564		160
19.08	601	190	60
0	609	-	-

Lithium chloride (LiCl) + Lead chloride ($PbCl_2$)

Treis, 1914

mol%	f.t.	E	min.
100	496	-	-
99.20	487	405	60
96.33	456	408	250
92.41	427	409	400
88.91	410	410	450
81.38	464	410	420
68.61	521	410	310
42.15	584	407	110
0	607	-	-

Bukhalov and Aleshkina, 1953 (fig.)

mol%	f.t.
100	498
64	400
43	490
12	562 I - II
0	606

Markov, Delimarskii and Panchenko, 1954 (fig.)

mol%	electromotive force 550°	600°
100	1.248	1.215
80	.256	.225
70	.267	.237
60	.267	.235
50	.272	.244
40	.280	.252
30	.288	-

Gromakov and Gromakova, 1955

mol %	f.t.	E
0	610	-
20	556	-
30	526	-
40	487	-
50	460	-
60	422	400
70	417	400
80	450	-
90	478	-
100	500	-

Lithium chloride (LiCl) + Manganese chloride ($MnCl_2$)

Sandonnini and Scarpa, 1913

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	650	-	50	562	555
90	630	605	40	575	563
80	615	578	20	586	578
70	583	559	10	595	585
60	563	555	0	602	-

Lithium chloride (LiCl) + Cobalt chloride ($CoCl_2$)

Ferrari and Baroni, 1928

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	602	-	52.0	538	535
6.3	593	589	53.0	533	530
14.5	589	586	56.0	529	523
21.5	580	578	60.0	530	524
25.0	572	569	62.5	533	530
32.0	560	555	64.5	542	539
35.5	555	553	68.1	564	562
39.8	548	546	72.9	592	590
43.7	546	545	76.5	625	624
45.4	549	547	82.0	650	642
47.5	553	548	88.2	690	689
48.7	556	548	92.4	720	716
50.1	558	554	100	773	-
51.0	554	551			

Lesnikh and Bergman, 1953

mol%	f.t.	mol%	f.t.
100	732	31	519
90	707	30	522
82	680	25	530
74	656	21	539
67	628	18	548
60	600	14	557
54	572	11	565
48	549	8	575
43	532	5	585
38	521	3	595
35	518	0	608
33	517 E		

Lithium chloride (LiCl) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

%	f.t.	%	f.t.
100.0	190.2	56.8	117.8
93.4	190.0	54.2	132.8
82.8	187.9	52.6	139.9
79.9	186.9	50.4	143.0
70.4	171.4	50.6	170.4
61.6	125.6	49.5	338.5
59.9	114.4		

Sodium chloride (NaCl) + Potassium chloride (KCl)

Heterogeneous eqilibria

Hackspill and Grandadam, 1924

b.t.	p	b.t.	p
100%			
830	1	920	5
868	2.4	940	6.2
890	3.5	970	8.1
		994	10.8

%	b.t.		
	1.08 mm	2.44 mm	3.49 mm
100	830	870	890
90	836	874	896
80	841	881	902
70	848	888	910
60	857	896	917
50	864	903	925
40	873	911	933
30	883	919	941
20	894	927	950
10	907	936	958
0	920	944	965

%	dew point (3.49 mm)	%	dew point (3.49 mm)
100	890	40	940
90	908	30	947
80	917	20	952
70	923	10	959
60	930	0	965
50	934		

Kordes and Raaz, 1929

wt%	mol%	b.t.	wt%	mol%	b.t.
100	100	1411	56.0	50	1437
87.9	85	1425	35.2	30	1437
71.8	70	1435	12.4	10	1433
65.6	60	1435	0	0	1430

Greiner and Jellinek, 1933

mol% L	mol% V	p_1	p_2
1180°			
0	0	73	0
25	31	57.1	26
50	55.6	36.7	46.3
75	82.8	20.3	76.3
100	100	0	106

Kangro and Wieking, 1938

mol%	p	p_1	p_2
860°			
25.2	2.9 ± 0.5	1.8	1.1
49.8	3.1 ± 0.5	1.0	2.1
74.7	4.0 ± 0.4	0.7	3.3
950°			
25.8	9.6 ± 0.8	6.5	3.1
49.6	9.7 ± 0.7	3.5	6.2
74.4	10.8 ± 0.8	1.9	8.9
990°			
24.3	16.5 ± 0.7	9.8	6.7
49.5	16.5 ± 0.7	7.3	9.2
75.9	17.9 ± 0.7	4.9	13.0

Le Chatelier, 1894

mol%	f.t.	mol%	f.t.
0	780	50	640
42	660	74	690
45	650	100	740

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	790	50	675
90	760	40	685
80	735	34	710
70	700	0	820
60	685		

Kurnakov and Zhemchuzhni, 1907

mol%	f.t.	tr.t.	min.
0.00	819	-	-
5.29	804	-	-
13.56	775	-	-
20.00	740	373-386	0.5
25.00	720	395	3.5
33.00	694	406	7.8
41.70	670	405	7.5
50.00	664	395	6.5
54.68	665	386	6.5
63.54	679	387	4.5
70.70	700	320-331	0.5
86.25	748	-	-
100.00	790	-	-

Jänecke, 1908

mol%	f.t.	sat.t. of mixed crystals
50	670	380-320
80	745-710	360

Benrath and Wainoff, 1911

%	f.t.	m.t.	%	f.t.	m.t.
100	785	785	40	768	675
90	786	760	30	770	700
80	786	725	20	780	740
70	777	695	10	740	780
60	770	675	0	819	819
50	764	670			

Gemsky, 1913

%	f.t.	sat.t. of mixed crystals
0	798	-
18.37	760	351
35.34	696	360
56.05	660	373
63.31	664	371
83.61	700	343
100	775	-

Treis, 1914

%	f.t.	m.t.	sat.t. of mixed crystals
100	775	775	-
84.9	737	692	-
75.0	710	668	340
60.0	673	663	382
49.9	663	661	386
19.7	724	673	339
0	798	798	-

Lantsberry and Page, 1920

%	f.t.	%	f.t.
0	785	60	652
10	760	70	662
20	735	80	690
30	700	90	715
40	667	100	760
50	652		

Akopov and Bergman, 1954

mol%	f.t.	mol%	f.t.
100	774	40	665
85	760	25	700
70	697	0	800
50	660		

Bukhalova and Bergman, 1955

%	f.t.	%	f.t.
0	800	60	670
10	772	75	720
25	722	100	775
35	676		

Rubleva and Bergman, 1956.

mol%	f.t.	mol%	f.t.
0	770		
10	766	40	666
15	749	45	661
20	732	50	658
25	716	55	661
30	696	60	669
35	680	65	681

tr.t.: 550 °

Nacken, 1918

t	mol% mixed crystals C_1	C_2
465	56	15
400	74	94
335	88	98

Bunk and Tichelaar, 1953

t	mol% C_1	C_2
309	88.9	2.1
391	77.1	6.1
447	63.4	13.7
466	54.2	19.5
472	52.1	19.1
488	33.5	-

Barret and Wallace, 1954

t	mol% C_1	C_2	t	mol% C_1	C_2
367	83	4.4	462	61.2	93.9
367	83	4.4	462	61.2	13.8
422	70.9	9.6	496	43.6	29.2
422	71.2	9.6			

Matsen and Beach, 1941

Mecanism and rate of formation of solid solutions in equimolar mixtures at several temperatures (by X-ray powder photographs)

Properties of phases

Van Artsdalen and Yaffe, 1955

t	d	t	d
	0%		
803.1	1.5560	898.9	1.5031
816.4	.5467	916.9	.4928
819.7	.5457	918.2	.4921
826.0	.5425	922.5	.4909
826.1	.5424	939.9	.4812
829.3	.5401	940.8	.4798
830.0	.5397	942.2	.4793
840.3	.5362	953.6	.4743
840.6	.5361	967.2	.4661
845.1	.5317	967.6	.4653
859.2	.5242	986.0	.4566
866.6	.5200	995.8	.4494
867.2	.5198	1004.9	.4468
870.9	.5184	1014.1	.4403
889.2	.5086	1027.5	.4322
890.3	.5074		
890.4	.5073		
891.6	.5071		
	15.23%		
780.6	1.5568	850.1	1.5178
781.3	.5564	875.3	.5037
807.1	.5414	898.3	.4926
829.8	.5288	921.2	.4787
849.9	.5180		
	27.06%		
707.5	1.5895	823.2	1.5241
739.6	.5711	859.4	.5041
777.9	.5504	893.7	.4854
802.3	.5358	915.6	.4730
	34.85%		
705.5	1.5886	823.0	1.5229
758.5	.5587	855.8	.5042
779.0	.5586	885.1	.4883
803.5	.5333	886.2	.4877
818.4	.5251	924.7	.4662
	48.77%		
670.0	1.5956	817.5	1.5118
710.9	.5728	842.2	.4981
752.1	.5493	871.5	.4814
794.6	.5250	909.0	.4600
	50.00%		
685.0	1.5837	796.2	1.5195
685.1	.5836	826.0	.5022
704.1	.5729	838.1	.4955
740.7	.5514	875.6	.4738
748.6	.5465	908.2	.4556
766.5	.5263		
783.4	.5267		
	79.25%		
717.0	1.5622	855.0	1.4819
738.6	.5498	855.4	.4816
758.8	.5380	878.2	.4685
759.0	.5398	878.4	.4683
783.4	.5235	905.7	.4525
801.8	.5130	927.8	.4397
802.3	.5127	928.2	.4395
835.1	.4932		
	100%		
779.5	1.5219	870.4	1.4688
781.8	.5211	907.2	.4479
809.4	.5058	939.0	.4292
827.4	.4943		

Sandonnini, 1920

%	d	%	d
	850°		
100	1.465	35.0	.490
95.0	.466	25.0	.494
86.58	.470	10.0	.504
75.0	.477	0	.512
50.0	.484		

Barzakovskii, 1940

mol%	molar volume* (in cc) 800°	900°
0	37.78	39.05
25	40.86	42.36
50	43.78	45.46
75	46.71	48.44
100	49.73	51.49

*calculated from density measurements of Mashovets

mol%	η 800°	900°
0	1590	1000
50	1170	900
75	1070	810
100	1130	890

mol%	σ*	mol%	σ
	800°		
0	114.7	70.0	102.8
24.6	110.2	83.6	98.9
46.6	105.6	100	96.1
56.2	104.2		

*data taken from Jirov, 1935

Desyatnikov, 1956.

mol%	σ 800°	900°	dσ /dt
100	-	89	-
90	97.6	90	0.075
75	98.6	91	0.074
60	100.0	93	0.072
50	102.0	94	0.072
40	103.2	95	0.072
25	106.1	98	0.070
10	110.5	103	0.068
0	-	107	-

Nacken, 1918

%	n_D	%	n_D
0	1.5443	60	1.5093
10	.5382	70	.5048
20	.5316	80	.5000
30	.5258	90	.4955
40	.5205	100	.4903
50	.5145		

Bunk and Tichelaar, 1953

mol %	lattice const. (in Å)	mol %	lattice const. (in Å)
100	6.28	32.89	5.865
87.58	6.205	25.17	5.81
75.80	6.14	16.40	5.75
54.05	6.015	8.02	5.69
43.95	5.945	0	5.62

Arnolt and Geszler, 1908

t	κ 100 %	50 %	0 %
800	2.19	2.71	3.34
850	2.30	2.86	3.50
900	2.40	2.98	3.66
950	2.50	3.07	3.82

Benrath and Wainoff, 1911

t	κ.10⁶ 100 %	t	100 %
640	5.0	720	36.9
650	6.5	730	43.7
660	8.4	740	59.6
670	12.0	750	77.8
680	13.4	760	91.9
690	18.6	770	133.8
700	20.9	780	199.2
710	28.8	790	1908000

t	90 %	80 %	70 %	60 %	50 %
480	1.5	4.2	5.4	4.9	6.0
490	1.7	5.6	5.7	6.0	8.0
500	2.6	6.0	6.0	7.1	10.0
510	3.0	7.0	7.5	9.0	10.5
520	3.5	8.0	9.0	11.5	11.5
530	4.5	9.7	12.0	12.2	13.0
540	5.1	11.2	13.1	14.7	14.5
550	5.8	13.4	15.3	17.0	15.0
560	6.8	15.4	18.4	19.5	20.0
570	8.0	16.5	22.0	24.0	24.0
580	9.0	20.2	27.0	30.0	32.0
590	10.8	23.8	31.5	38.0	40.0
600	12.5	27.0	37.0	47.0	52.0
610	14.5	31.0	44.0	57.0	64.0
620	17.0	36.0	54.0	70.5	80.0
630	20.0	42.6	66.0	89.0	100.0
640	23.5	50.0	81.0	111	121.0
650	27.8	60.0	101	150	154.0
660	33.0	70.0	123	170	204
670	40.0	88.0	160	13500	8040
680	50.0	120	25000	-	-
690	100000	6620	-	-	-

t	40%	30%	20%	10%
480	7.0	7.0	4.0	3.0
490	8.0	8.0	8.0	5.0
500	10.5	11.0	10.0	7.0
510	12.0	12.5	13.5	10.0
520	14.0	14.5	16.0	12.0
530	16.0	17.0	20.0	14.0
540	20.0	21.0	25.0	16.0
550	24.0	25.0	30.0	20.0
560	29.0	30.0	36.0	24.5
570	30.0	34.5	40.0	28.0
580	35.0	42.0	48.0	35.0
590	44.0	51.0	54.0	40.0
600	53.0	64.0	62.0	46.0
610	68.0	80.0	72.0	55.0
620	88.0	101	82.0	60.0
630	116.0	126	96.0	68.0
640	156.0	160	112	79.0
650	210	220	135	90.0
660	-	-	164	102
670	-	25000	212	110
680	-	-	4950	135
690	-	-	-	155
700	-	-	-	177
710	-	-	-	210
720	-	-	-	235
730	-	-	-	5400
740				

t	0%	t	0%
540	2.0	690	60.0
550	2.8	700	75.0
560	3.8	710	95.0
570	4.5	720	120
580	5.0	730	150
590	6.1	740	190
600	8.4	750	240
610	9.7	760	300
620	11.0	770	360
630	12.6	780	425
640	17.0	790	500
650	21.2	800	729
660	28.1	810	1500
670	35.0	819	3339000
680	45.0		

Ryschkewitsch, 1933

t	κ (in mhos)	t	κ	t	κ	t	κ
0 mol%		10 mol%		20 mol%		50 mol%	
805	3.54	805	3.33	820	3.23	720	2.39
820	3.62	813	3.36	824	3.22	730	2.46
833	3.68	823	3.38	832	3.26	740	2.49
80 mol%		100 mol%				748	2.53
						760	2.58
785	2.57	805	2.45			810	2.82
788	2.60	818	2.49			830	2.97
795	2.66	825	2.51				
800	2.67	840	2.67				
815	2.77	850	2.67				
		855	2.60				

Van Artsdalen and Yappe, 1955

t	κ (in mhos)t		κ
	0%		
802.3	3.0887	899.4	3.8637
804.6	.4497	911.5	3.9154
806.4	.5804	913.3	.9129
806.5	.5821	922.4	.9416
808.0	.6057	933.3	.9680
808.2	.6060	933.8	.9701
817.3	.6431	947.3	4.0151
817.9	.6441	971.5	.0653
827.4	.6601	972.0	.0616
839.8	.7145	972.3	.0719
857.8	.7662	1006.9	.1595
871.4	.8041	1010.9	.1701
892.7	.8502	1021.4	.1861
	15.23%		
758.0	2.9859	816.8	3.2787
759.6	3.0627	817.3	.2803
762.7	.1085	835.0	.3309
767.2	.1237	852.4	.3784
773.2	.1444	875.3	.4415
789.4	.1938	895.5	.4885
795.2	.2092	909.2	.5233
811.2	.2607	927.2	.5609
	27.06%		
708.0	2.5583	819.5	3.0365
717.1	.6866	841.0	.0947
735.1	.7591	857.7	.1397
750.5	.7591	876.4	.1914
773.9	.8977	899.4	.2344
788.8	.9472	927.9	.2864
	34.85%		
690.7	2.3236	825.0	2.9254
695.4	.5049	853.4	.9281
709.1	.5562	898.4	3.0066
736.7	.6556	919.7	.1611
756.8	.7175	944.2	.2163
783.6	.7971		
823.6	.9210		
	48.77%		
665.9	2.1986	788.9	2.6107
674.4	.2523	818.8	.6914
679.2	.2789	843.4	.7523
706.2	.3687	846.7	.7569
726.8	.4374	875.1	.8273
738.7	.4695	887.6	.8533
769.5	.5633	912.5	.9019
	59.00%		
673.5	1.5316	769.8	2.4472
679.0	.9085	794.5	.5196
692.1	2.1766	822.8	.5848
708.3	.2370	853.0	.6641
734.9	.3393	853.7	.6641
740.7	.3639	881.5	.7223
740.9	.3643	907.9	.7725
	79.60%		
713.4	1.8246	815.1	2.3784
718.9	.8344	839.0	.4379
730.5	2.1446	853.7	.4638
730.9	.1553	868.4	.4958
766.4	.2466	889.8	.5422
789.7	.3180	930.1	.6109
	100%		
778.9	2.1628	850.9	2.3609
782.4	.1816	851.2	.3618
787.0	.1984	865.0	.3901
790.0	.2074	877.5	.4174
799.0	.2342	889.0	.4476
800.5	.2407	892.5	.4484
804.1	.2478	905.5	.4745
823.0	.3020	906.7	.4777
846.0	.3507	925.1	.5156

Mulcahy and Heymann, 1943 (fig.)

mol %	λ	mol %	λ
	850°		
100	121	30	126
90	120	20	130
70	120	5	136
45	125	0	138

Thermal properties .

Zhemchuzhni and Rambach, 1910

mol %	Q mix	mol %	Q mix
12	458	62	784
25	590	75	580
40	920	88	391
50	1046		

Bunk and Tichelaar, 1953

mol %	Q mix	mol %	Q mix
90	390	40	1060
80	708	30	958
70	931	20	750
60	1053	10	469
50	1686	0	0

Hyvönen, 1952

Evolution of heat during the ageing, dQ/dh (millicals/hours, mole)= H

50 mol%

h	H 1(24.93°)	2(24.95°)	3(24.95°)
5	97.9	-	-
6	96.1	261.4	326.5
24	10.9	36.3	91.8
120	3.1	5.1	-
168	-	1.4	-

h	H	h	H	h	H
		4			
25.12°		30.19°		35.19°	
5	86.6	97	68.8	145	143.6
88	11.4	140	14.3	164	23.3
40.18°		45.07°		49.80°	
169	155.9	194	72.1	218	72.9
188	27.9	210	40.2	260	50.3
		5			
25.04°		35.10°		45.13°	
3	97.6	101	309.2	150	63.0
44	27.6	140	49.1	212	64.1

h	H 6(24.93°)	7(24.93°)	8(25.08°)	9(35.18°)
3	417.7	-	-	-
5	402.7	-	-	330.1
6	400.2	572.8	-	316.4
11	386.5	559.5	426.3	279.7
40	324.0	519.8	392.8	211.4
64	298.9	-	374.2	191.3
192	-	-	307.7	156.0
312	-	-	310.8	-
	10(45.08°)	11(45.08°)	12(45.07°)	
3	-	574.0	371.9	
4	309.0	434.1	328.0	
96	32.0	280.4	34.5	
144	-	267.3	24.4	
192	-	244.7	-	

h= hours counted from the quenching; nos. 1,2,3 were annealed continuously for 20 h. after the crystallisation. nos 4 and 5 were at first cooled to ro om temperature and then annealed for 48 h.; 6-12 were rapidly cooled

Sodium chloride (NaCl) + Rubidium chloride (RbCl)

Kangro and Wieking, 1938

mol%	p	p_1	p_2
	860°		
26.2	2.5	2.0	0.5
50.6	2.4	1.3	1.1
75.3	2.4	0.6	1.8
	950°		
25.7	8.4	6.4	2.0
50.7	8.2	4.5	3.7
75.6	8.2	2.2	6.0
	990°		
26.0	15.0	11.0	4.0
50.8	14.4	7.4	7.0
75.7	14.3	4.0	10.3

Zhemchuzhni and Rambach, 1910

mol%	f.t.	E	mol%	f.t.	E
100	726	-	50.00	564	541
95	703	541	45.85	583	"
84.40	653	"	40.31	620	"
77.15	630	"	35.00	641	"
70.70	600	"	29.00	678	"
65.32	580	"	23.95	709	"
60.40	548	"	19.10	736	"
56.84	541	"	11.80	758	"
55.96	541	"	4.83	798	540
53.00	547	"	1.70	811.5	540
			0	819	-

Sodium chloride (NaCl) + Cesium chloride (CsCl)

Zhemchuzhni and Rambach, 1910

mol%	f.t.	E	tr.t.
100	646	-	451
96.79	632	485	451
94.99	619	487	451
87.76	584	490	451
81.85	560	493	451
75.56	535	493	451
72.00	516	493	451
66.00	496	493	451
65.50	493	493	451
60.00	516	493	451
50.00	580	493	451
40.00	638	493	451
28.00	701	493	451
12.00	772	493	451
6.00	792.5	490	450
2.84	806.5	490	447
0	819	-	-

Kangro and Wieking, 1938

mol%	p	p_1	p_2
860°			
26.8	3.2	2.0	1.2
47.4	3.9	1.2	2.7
72.7	5.3	0.5	4.8
950°			
26.7	10.4	5.8	4.6
48.7	12.8	3.2	9.6
72.6	17.6	1.8	15.8
990°			
27.7	18.1	12.0	6.1
45.8	22.9	7.1	15.8
69.4	31.8	3.6	28.2

Sodium chloride (NaCl) + Thallium chloride (TlCl)

Sandonnini and Aureggi,1911

mol%	f.t.	E	min.
0	806	-	-
10	770	410	30
25	711	410	-
40	667	412	50
50	650	415	-
65	570	412	60
70	530	414	-
75	490	412	70
80	430	-	-
85	-	412	100
90	423	412	50
100	429	-	-

Sodium chloride (NACl) + Cuprous chloride (CuCl)

Sandonnini, 1911

mol%	f.t.	E	min.
0	806	-	-
10	772	-	-
20	736	320	-
30	674	318	60
40	610	318	80
50	555	318	90
60	460	317	110
65	416	316	150
70	373	314	200
75	E	314	230
77.50	-	316	110
80	334	316	60
85	360	316	-
90	374 - 352	-	-
95	390 - 370	-	-
100	422	-	-

De Cesaris, 1911

%	f.t.	E	min.
0	796	-	-
10	875	-	-
20	755	322	120
30	723	324	195
40	680	324	315
50	610	321	390
60	545	320	450
70	440	324	525
80	-	325	555
86	350	324	300
90	370	322	60
95	395	-	-
100	420	-	-

Korreng, 1914

%	f.t.	E	min.
100	425	-	-
99.25	415	292	30
98.47	414	296	90
96.82	399	311	150
93.13	380	325	300
88.77	350	325	630
86.29	336	326	750
83.56	334	324	820
80.55	375	325	770
77.21	415	324	710
73.49	464	326	630
69.31	509	326	580
59.22	593	325	450
45.85	666	324	300
29.35	739	323	150
15.13	774	321	80
0	800	-	-

Sodium chloride (NaCl) + Silver chloride (AgCl)

Sandonnini, 1911

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	806	-	70	565	476
10	770	730	75	530	470
20	748	680	80	520	465
30	708	640	85	505	461
40	670	580	90	490	460
50	640	520	95	470	457
60	610	490	100	455	-
65	580	482			

Botta, 1911

%	f.t.	%	f.t.
0	792	71.02	652
21.40	752.5	78.62	613
37.99	742	85.12	592
59.23	712	90.74	544
62.03	692	95.66	510
		100	460

Jänecke, 1915

mol%	sat.t. of mixed crystals
90	near 98
75	131
50	157
25	148
10	84

Zhemchuzhni, 1916

mol %	f.t.	m.t.	mol %	f.t.	m.t.
100	451	-	47.7	667	520
96.6	458	452	43.7	674	543
92.0	478	457	39.7	698	572
86.40	507	460	35.2	720	600
81.5	534	462	25	750	660
75.8	560	465	17	772	710
70.20	579	469	10	788	737
64.5	610	473	6.9	795	762
53.8	646	485	3.1	809	797
51.9	655	500	0	816	-

Tubandt and Abramovitsch, 1927 (fig.)

mol %	f.t.	m.t.	mol %	f.t.	m.t.
100	440	440	40	695	595
90	500	450	20	760	650
80	540	475	0	820	820
60	625	500			

		κ		
mol%	300°	350°	400°	440°
100	18.5	70.0	-	-
90	10.0	25.0	-	-
80	4.0	20.0	88	-
70	2.0	10.0	40	110
60	0.5	6.0	17	30
50	-	1.0	50	11
40	-	-	10	5
30	-	-	-	1.5
20	-	-	-	-
84	-	-	120	-

Tubandt, Reinhold and Jost, 1927

280°

mol %	κ	mol %	κ
0	0.1	90	5.00
25	0.04	95	8.50
50	0.14	100	11.00
75	1.70		

Kolotii, 1956

mol %			e		
	830°	840°	850°	860°	870°
0.23	0.4480	0.4590	0.4700	-	-
1.65	.3313	.3359	.3387	0.3426	0.3470
4.75	.2445	.2473	.2504	.2526	.2559
33.39	.0785	.0792	.0803	.0813	.0822

Sodium chloride (NaCl) + Beryllium chloride ($BeCl_2$)

Schmidt, 1929

mol%	f.t.	m.t.	min.
0	775	-	-
10	-	305	4
20	562	310	8
30	400	310	18
35	-	310	14
40	298	212	12
45	266	213	15
50	224	215	20
55	236	215	18
60	270	215	15
70	320	214	12
80	355	205	6
90	375	210	6
100	404	-	-

E 51 mol% 215°

Delimarskii, Cheiko and Feshchenko, 1955

%				κ			
	260°	280°	300°	320°	340°	360°	380°
35.09	-	-	-	-	-	8500	9440
37.77	-	-	-	-	7300	8480	9340
38.43	-	-	-	-	7850	8740	9470
41.43	-	-	6900	7750	-	-	-
41.53	-	-	-	-	8650	9530	10340
43.19	-	5900	6740	7600	8470	9360	10080
44.86	-	5550	6490	7440	8370	9260	9890
50.74	4210	5080	5850	6710	7570	8360	9210
51.63	4110	4890	5720	6510	7360	8180	8950
54.92	3010	3760	4400	5190	5940	6680	7340
56.17	-	4120	4870	5700	6510	7330	7970
57.68	-	3240	3920	4790	5450	6250	7040
64.1	-	-	-	3720	4230	4710	5280
	400°	420°	440°	460°	480°	500°	520°
30.44	-	-	-	12440	12940	13540	14090
31.95	-	-	11390	12180	12820	13460	-
35.09	11190	10900	11690	12320	12820	13300	-
37.77	9950	10630	11330	11880	12350	12840	-
38.43	10160	10860	11620	12150	12690	13260	-
41.53	10990	11670	12330	12900	13360	13790	-
43.19	10670	11340	11740	12100	12460	-	-
44.86	10520	11160	11790	12360	12750	-	-
50.74	9790	10560	11310	11880	12380	-	-
51.63	9500	10260	10860	11450	11960	12550	-
54.92	7820	8340	8860	9480	9960	10480	-
56.17	8500	9160	-	-	-	-	-
57.68	7610	-	-	-	-	-	-
64.1	5650	6160	6680	-	-	-	-
78.48	1910	2290	2670	3030	3440	-	-

Sodium chloride (NaCl) + Magnesium chloride ($MgCl_2$)

Menge, 1911

%	f.t.	tr.t.	min.	E	min.
100	711	-	-	-	-
93.1	674	446	7.8	419	3.6
88.5	650	449	12.8	429	5.1
74.3	579	452	13.3	432	9.8
63.9	490	445	15.3	429	19.6
58.9	439	-	20.0	430	25.8
56.1	-	-	-	430	48.8
50.7	438	-	-	430	36.8
47.5	492	-	-	437	37.7
42.4	545	460	1.4	431	30.8
33.2	638	468	5.3	429	26.3
23.0	712	466	6.3	430	18.7
15.6	757	460	5.9	429	9.6
8	782	460	3.8	423	5.1
0	803	-	-	-	-

Scholich, 1919

mol %	wt %	f.t.	E	min.	tr.t.	min.
0	0	798	-	-	-	-
20	28.94	713	441	100	476	60
30	41.12	609	441	350	477	70
40	52.07	494	444	420	476	40
45	56.85	463	443	480	-	-
50	61.97	451	444	620	-	-
55	66.57	456	445	340	-	-
60	70.96	528	443	150	461	200
70	79.18	599	443	100	462	200
80	86.70	650	441	60	462	140
100	100	711	-	-	-	-

Speranskaya, 1938

mol %	f.t.	tr.t.	E
100	712	-	-
71.4	618	464	440
60.0	553	460	435
53.2	505	463	442
52.7	496	-	-
51.3	480	-	-
49.9	463	-	-
48.8	461	-	-
48.4	460	-	440
47.4	455	-	-
46.3	448	-	-
45.5	446	-	-
45.1	445	-	-
43.5	442	-	-
42.9	441	-	-
42.1	441	-	-
40.7	444	-	-
39.5	454	-	-
38.2	467	-	442
38.1	468	-	-
37.0	480	-	-
36.8	509	-	440
35.6	504	-	-
24.6	642	470	440
14.3	732	476	441

E_1 : 49.8 mol % 462° (1+1)
E_2 : 42.4 mol % 442° (1+2)
E_3 : 29.9 mol % 476°

Klemm and Weiss, 1940

mol%	f.t.	mol%	f.t.
100	720	36	485 = E_3
90	700	34	540
80	675	30	590
70	700	25	645
60	590	20	680
55	560	15	730
50	515 = E_1	10	750
45	462	0	795
40	450 = E_2	(1+1)	(1+2)

Desyatnikov, 1956

mol%	σ 700°	σ 800°	σ 900°	dt/d
0	121	-	107	-
10	-	103.8	98	0.056
20	101	96.6	92	0.043
30	96	92.5	88	0.035
40	93	89.0	85	0.033
50	89	85.7	82	0.027
60	84	80.8	78	0.023
70	80	77.0	75	0.020
80	75	73.1	71	0.014
90	71	69.0	67	0.012
100	67	-	65	-

Markov, Delimarskii and Panchenko, 1955

mol %	e	τ (10^3)	mol %	e	τ (10^3)
		718°			
100	2.531	-0.69	32.3	2.628	-0.576
55.9	.558	-0.565	24.2	.669	-0.615
52.1	.564	-0.55	22.2	.679	-0.566
40.3	.600	-0.63	23.6	.684	-0.572
37.9	.614	-0.507			

e = electromotive force, in volts .

Sodium chloride (NaCl) + Calcium chloride ($CaCl_2$)

Menge, 1911

wt%	f.t.	tr.t.	min.	E	min.
100	777	-	-	-	-
90	695	-	-	498	13.4
80	627	-	-	500	23.6
75	579	-	-	503	28.0
60	507	-	-	501	32.3
68	-	-	-	500	37.5
65.5	527	-	-	503	24.0
63	556	-	-	505	32.1
60	577	-	-	497	27.0
57	594	-	-	499	25.5
54	610	-	-	503	22.2
50	638	572	1	502	14.0
40	692	607	1	500	6.3
30	732	591	3.5	494	3.6
20	764	565	3	493	1.8
10	795	-	-	-	-
0	808	-	-	-	-

Lamplough, 1912 (fig.)

mol%	f.t.	mol%	f.t.
100	767	40	595
80	685	20	725
60	565	0	800
53	495 E		

Scholich, 1919

mol%	wt%	f.t.	E	min.
0	-	798	-	-
2	3.73	795	511	-
10	17.42	748	501	15
20	32.19	714	509	40
30	44.87	668	502	90
40	55.87	602	502	170
50	65.60	518	508	260
60	74.01	548	506	250
80	88.37	678	506	130
100	-	770	-	-

Lantsberry and Page, 1920

%	f.t.	%	f.t.
100	795	70	508
90	657	60	570
80	540	50	620
75	508	25	740
72.5	505	0	785

Gromakov and Gromakova, 1955

mol%	f.t.	E	mol%	f.t.	E
0	800	800	55	511	490
9.5	761	-	60	552	-
17	668	490	70	620	490
30	653	490	80	676	490
40	587	490	90	719	600
45	550	490	100	770	-
50	(510)	490			

Sandonnini, 1920

%	d	κ(in mhos)	%	d	κ
		850°			
100	2.220	2.030	57.74	2.520	1.766
97.49	.191	.005	50.0	2.635	.726
94.86	.190	1.988	40.0	2.830	.678
90.0	.225	.956	25.0	3.016	.612
80.0	.307	.892	20.0	3.260	.587
64.5	.404	.786	0	3.575	.512
60.0	.452	.77			

Borzakovski , 1940

mol%	d	$\tau.10^5$	mol%	d	$\tau.10^5$
		800°			
0	1.544	60	52.81	1.831	46
10.23	.607	55	62.9	.886	42
19.6	.667	54	82.9	.967	44
43.48	.781	47	100	2.049	48

			η		
mol%	600°	700°	800°	900°	1000°
0	-	-	1590	1000	700
10	-	-	1650	1160	1000
30	-	3180	2590	2110	1790
40	-	3820	3110	2590	2310
50	5400	4300	3490	2950	2600
60	6000	4800	3790	3180	2830
80	-	-	4360	3690	3310
90	-	-	4630	4000	3530
100	-	-	4920	4220	3740

mol%	σ cm/dyne	
	800°	1000°
0	114.0	99.6
10	115.1	100.0
30	121.2	108.0
50	126.3	115.8
80	137.2	125.1
100	147.6	140.0

Ryschkewitsch, 1933

t	κ (in mhos)	t	κ	t	κ
0%		33%		50%	
830	3.35	700	2.14	655	1.505
840	3.40	750	2.82	740	1.99
880	3.525	840	3.175	840	2.62
930	3.735	890	3.42	860	2.74
940	3.965	960	3.80	890	2.82
960	4.02			960	3.31
67%		100%			
650	1.27	820	2.22		
690	1.49	840	2.34		
755	1.87	870	2.56		
790	2.16	920	2.74		
830	2.34	950	2.87		
850	2.49				
860	2.64				
900	2.82				
950	3.01				

Borzakovskii, 1940

mol%	κ (in mhos) 600°	700°	800°	900°
0	-	-	3.57	3.79
12.5	-	2.70	3.08	3.25
25	-	2.25	2.65	2.97
30.9	-	2.115	2.51	2.855
45	1.465	1.905	2.28	2.605
62	1.23	1.695	2.09	2.45
75.5	-	1.56	1.95	2.305
88.75	-	-	1.98	2.31
100	-	-	2.02	2.33

Sodium chloride (NaCl) + Strontium chloride ($SrCl_2$)

Vortisch, 1914

%	f.t.	E	min.
0	798	-	-
23.16	765	564	80
40.41	721	565	160
53.75	670	565	200
64.39	623	564	270
73.06	-	565	350
80.27	620	564	250
86.35	685	565	160
91.56	742	564	110
96.06	805	559	60
100	870	-	-

Sodium chloride (NaCl) + Barium chloride ($BaCl_2$)

Le Chatelier, 1894 and 1897

mol%	f.t.	mol%	f.t.
0	778	60	645
10	758	70	695
20	740	80	750
30	717	87	770
40	690	90	830
50	640	100	910
56	615		

Sackur, 1912

mol/1000g NaCl	f.t.
0	802
0.41	793
.46	791
.76	786
1.19	777

Gemsky, 1913

%	f.t.	E	min.	tr.t.	min.
0	798	-	-	-	-
28.36	765	654	80	-	-
47.11	730	655	150	-	-
54.29	710	654	190	-	-
60.43	689	654	230	-	-
65.74	669	654	270	-	-
70.37	-	654	410	-	-
78.08	715	654	240	-	-
84.24	768	654	220	-	-
89.26	816	653	160	-	-
93.44	863	653	120	-	-
94.56	876	653	90	-	-
96.98	901	649	60	-	-
99.43	935	649	20	931	10
100	955	-	-	930	80

Vortisch, 1914

%	f.t.	E	min.	tr.t.	min.
0	798	-	-	-	-
54.29	708	654	180	-	-
78.08	710	654	200	-	-
91.44	834	652	120	-	-
99.43	934	646	20	922	80
100	955	-	-	922	80

Gromakov and Gromakova, 1955

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	800	800	40	-	648
10	772	742	50	707	648
14	746	662	60	765	648
30.5	715	648	100	-	960

Barzakovskii, 1940

mol%	molar volume (in cc) 800°	1000°
0	37.78	41.04
20	42.2	44.7
35	47.0	49.5
50	49.4	52.6
65	54.5	57.3
100	-	67.1

mol%	η 900°	1000°
0	1000	700
10	800	510
25	1400	1120
50	2827	2290
75	4190	3360
100	-	5110

mol%	σ 900°	1000°
0	107.0	99.6
10	108.6	103.3
20	117.0	110.9
35	126.1	119.9
50	132.9	126.4
80	159.9	153.7
90	-	162.2
100	-	171.0

mol%	$\varkappa$ 900°	1000°
0	3775	4010
13.0	3110	3340
23.95	3750	3020
35.4	2590	2845
48.0	2340	2655
67.5	2210	2550
83.5	1940	2265
100	-	2045

Sodium chloride (NaCl) + Stannous chloride ($SnCl_2$)

Rack, 1913

%	f.t.	E	min.
100	239	-	-
99.47	237	176	-
98.41	235	178	40
96.69	231	181	160
94.85	225	181	240
92.85	210	182	360
88.35	187	183	420
86.66	250	183	520
82.97	398	183	460
76.46	510	183	400
61.90	673	184	340
44.83	721	182	240
0	800	-	-

Sodium chloride (NaCl) + Lead chloride ($PbCl_2$)

Treis, 1914

%	f.t.	E	min.
100	496	-	-
97.72	477	407	180
95.01	442	409	310
92.78	414	408	460
91.73	413	409	540
89.82	447	410	520
87.71	471	412	520
82.62	541	411	500
76.02	608	411	460
67.09	679	411	420
54.32	727	411	340
34.57	767	408	220
0	798	-	-

Gromakov and Gromakova, 1955

mol%	f.t.	mol%	f.t.
0	800	58	480
20	706	60	453
26	675	64	438
30	654	66	424
35	623	68	406
40	594	70	379
42	584	71.5	401
44	575	74	412
46	560	76	422
48	545	78	427
50	534	80	433
52	516	90	495
54	506	100	496
56	492		

Barzakovskii, 1940

mol%	molar volume (in cc)* 500°	600°
100	56.13	57.70
88.4	53.45	55.29
77.2	52.58	54.07
66..9	48.77	50.33
57.8	46.87	49.20
49.8	45.82	48.01

* calculated from density measurements of Dvorkin and Artamonov

mol%	σ 500°	600°
0	142.8	131.4
10	139.8	129.0
25	135.9	125.1
50	137.0	127.1

mol%	ϰ 500°	600°
100	14450	19100
89	15100	19500
68	15800	20050
45.6	16450	20700

Markov, Delimarskii and Panchenko, 1954

mol%	electromotive force 550°	600°
100	1.248	1.215
80	.260	.230
70	.267	.233
60	.276	.247
55	.280	.249

Sodium chloride (NaCl) + Zinc chloride ($ZnCl_2$)

Pavlenko, Bergman and Nikonowa, 1941

mol%	f.t.	solid phase
100	320	2
94	316	"
90	315	"
85	313	"
80	308	"
75	303	"
72.5	297	"
70	293.5	"
65	280	"
60	268	"
58.5	262	E
57	281	(1+2)
55	294	"
52.5	312.5	"
50	325	"
45	358.5	"
52	382.5	"
40	393	"
37,5	403	"
36.5	410	transition
34	425	"
35	457	"
33	490	"
30	560	"
25	640	"
20	690	"
15	736	"
0	800	"

Gromakov and Gromakova, 1955

mol%	f.t.	mol%	f.t.	m.t.
0	800	70	243	-
30	520	80	263	250
40	395	90	273	-
50	350	100	280	-
60	290			

Sodium chloride (NaCl) + Cadmium chloride ($CdCl_2$)

Brand, 1911

%	f.t.	E	min.	tr.t. (2+1)	min.
100	562	-	-	-	-
96.58	536	395	50	-	-
92.62	506	394	100	-	-
87.97	463	394	190	-	-
82.47	415	393	270	-	-
79.31	-	392	380	-	-
75.82	407	390	280	-	-
74.32	412	391	230	-	-
67.64	483	388	110	423	100
62.80	549	390	40	426	200
61.06	576	-	-	425	210
57.34	602	-	-	423	190
50.00	658	-	-	419	170
43.86	690	-	-	411	150
25.84	756	-	-	402	90
14.17	780	-	-	391	50
0	798	-	-	-	-

Dergunov, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	562	45	433
80	510	40	490
60	405	20	705
57	382 E (2+1)		

Ilyasov and Bergman, 1956.

mol%	f.t.	mol%	f.t.
100	569	53.9	391
90.5	549	51.0	408
81.9	522	48.2	423
74	485	45.5	434
66.7	450	45.5	433 tr.t.
60	415	42.9	460
56.9	395	38	514
54.7	387 E	33.4	571

Bergman and Bakumskaya, 1955

mol%	f.t.	mol%	f.t.
100	568	48.2	422
90.5	546	45.5	432
81.9	518	42.9	460
74.0	486	40.4	490
66.7	452	38	520
63.3	432	33.4	572
60	410	29.1	618
56.9	394	25.0	656
53.9	394	0	800
50.5	408		

Boardman, Palmer and Heyman, 1955 (fig.)

mol%	σ 700°	σ 800°
0	121	115
20	108	100
40	98	91
60	90	85
80	85	81
100	81	79

Sodium chloride (NaCl) + Mercuric chloride ($HgCl_2$)

Belaiev and Mironov, 1952

mol%	f.t.	mol%	f.t.
100	283.0	84.8	268.0
95.0	279.0	83.0	268.5
90.0	271.5	82.0	276.0
95.5	270.0	81.1	292.5
87.0	267.5	80.0	303.0
85.5	267.5	78.5	309.0

E: 86% 264° tr.t. 272° 82.5%

Sodium chloride (NaCl) + Manganese chloride ($MnCl_2$)

Sandonnini and Scarpa, 1913

mol%	f.t.	$tr.t._1$	min.	$tr.t._2$	min.
100	650	-	-	-	-
90	628	441	-	-	-
80	594	443	40	-	-
70	556	442	70	-	-
65	520	442	45	427	-
60	490	441	-	426	-
55		442	-	425	40
50	-	-	-	425	10
45	448	-	-	424	60
40	460	446	40	425	-
35	520	445	60	424	-
30	580	445	70	424	-
20	600	445	70	-	-
10	740	442	20	-	-
0	804	-	-	-	-

Sodium chloride (NaCl) + Aluminum chloride ($AlCl_3$)

Plotnikov, Fortunatov and Galinker, 1933

mol%	p
700°	
45	155
29.4	94
25.75	38

Narishkin, 1939 (fig.)

t	p_2	t	p_2	t	p_2
65.6mol%		59.4mol%		51.8mol%	
182	80	265	16	431	3.5
153	100	308	32	473	10
253	316	393	100	547	32
329	1000	496	316	560	40
352	1410				

Morozov, Korshunov and Simonich, 1956.

t	p	t	p
50.0%		85.0%	
300	0.0	131	21.8
400	0.0	141	38.3
500	2.0	159	121.0
670	17.8		

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	190.5	58.9	123.6
79.9	192.0 compl.	55.2	140.7
77.9	190.0	52.4	147.9
73.9	182.0	51.7	151.9
69.4	169.4	51.6	150.7
66.1	151.3	50.5	320.5 (1+1)
52.7	130.2		

99.8-82.0 mol% L_1+L_2+C 193.5°

Shvartsman, 1940

%	f.t.
61	108 E
50	152 (1+1)
-	152 E

Chrétien and Lous, 1943(fig.)

%	f.t.	E	%	f.t.	E
85	135	105	60	635	155
81	105	105	40	775	155
75	150	105	20	790	155
69.5	155 (1+1)		0	800	-
68.8	455	155			

Kryagov, 1939

mol%	ϰ			
	190°	200°	210°	220°
50.3	4100	4360	4590	4800
50.7	-	3640	3920	4180
51.8	2300	2620	2900	3140
53.0	2540	2800	3060	3220
58.2	1220	1600	1800	-
59.4	-	1700	1800	2200
65.5	-	900	1180	1300
67.4	-	700	-	1100
	225°	230°	250°	270°
50.3	4900	5000	5320	5620
50.7	4300	4400	4820	5160
51.8	-	3330	3780	4100
53.0	3540	3660	3800	-
58.2	-	-	-	-
59.4	2450	2520	3000	3600
65.5	1600	-	1780	1900
67.4	-	1300	-	-

Dewing, 1955 (fig.)

t	-F			
	53.6	58.1	64.0	65.8
	mol%			
393	101.4	-	-	-
315	97.4	100.6	-	-
253	93.6	96.8	99.6	100.8
203	89.8	93	96.3	97.2
162	85.8	89.2	92.8	93.4
144	-	87.2	91	91.8

F=4.576 log. P+2.303DU log.T.where DU=-16cal/degr.

Sodium chloride (NaCl) + Indium chloride ($InCl_3$)

Vovkogon and Fialkov, 1945

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0.0	800	-	52.3	708	-
13.3	779	-	53.0	710	-
22.2	760	700	55.2	640	-
29.6	745	704	61.6	488	319
35.8	716	703	66.8	407	316
37.6	702	-	74.6	368	316
40.0	710	700	80.6	349	315
46.4	730	702	81.6	369	317
47.0	726	703	84.8	424	314
50.6	730	703(1+1)	94.6	546	321

E_1 : 78% 316° E_2 : 39% 704°

Sodium chloride (NaCl) + Zirconium chloride ($ZrCl_4$)

Belozerski and Kucherenko, 1942

wt %	mol %	f.t.	m.t.	wt %	mol %	f.t.	m.t.
0	0	800	-	56.3	24.5	500	-
8.7	2.2	680	408	57.1	25.0	470	395
10.4	2.6	603	380	57.7	25.7	480	406
13.2	3.2	560	390	58.9	26.6	400	-
22.3	6.2	515	405	62.3	29.0	395	-
22.8	6.4	-	395	70.0	37.0	270	220
22.8	6.4	515	380	72.0	39.0	250	220
27.0	8.2	480	390	74.6	42.0	217	-
28.3	8.8	465	388	76.3	44.8	307	-
35.4	12.2	423	395	77.0	45.2	305	220
37.3	13.3	460	372	80.0	50.0	330	230
37.6	13.4	488	400	86.2	60.4	217	167*
42.2	15.8	525	380	87.6	64.0	212	162
47.8	19.0	535	390*	88.8	66.0	262	160
50.9	20.8	530	390				

* (4 + 1) (1 + 1)

Morozov and Korshunov, 1956

%	t	p_2	%	t	p_2
35.0	247	0.0	48.0	353	0.7
35.0	276	0.0	48.0	412	4.9
35.0	311	0.3	48.0	423	8.2
35.0	361	0.9	48.0	487	25.5
35.0	362	1.5	72.0	246	18.3
35.0	437	10.1	72.0	273	18.1
35.0	493	11.4	72.0	305	132
35.0	590	50.0	72.0	310	183.5
48.0	282	0.0			

%	f.t.	E	tr.t. 1	tr.t. 2
100.0	437	-	-	-
95.0	404	316	-	-
90.0	318	313	-	-
88.0	315	315	-	-
85.0	318	312	-	-
80.0	356	310	-	339
77.0	373	308	-	338
75.0	378	306	375	343
72.0	483	310	374	341
70.0	576	308	369	350
68.0	631	-	370	351
65.0	688	515	376	340
62.0	701	540	377	341
60.0	690	541	377	341
55.0	639	542	377	340
52.0	540	-	378	340
50.0	651	538	375	340
48.0	692	538	375	348
40.0	717	549	383	340
38.0	720	539	375	346
35.0	737	540	370	344
32.0	726	536	371	335
30.0	744	535	370	335
25.0	756	532	377	338
20.0	768	536	376	341
15.0	775	539	-	-
10.0	786	-	-	-
5.0	788	-	-	-
0.0	800	-	-	-

Sodium chloride (NaCl) + Tantalum chloride ($TaCl_5$)

Morozov, Korchunov and Simonich, 1956

%	f.t.	tr.t. a*	tr.t. b*	%	f.t.	tr.t. a*	tr.t. b*
100.0	220	-	-	87.0	474	-	232
99.0	218	-	-	86.0	470	-	232
98.0	218	-	-	80.0	461	-	232
96.0	232	-	219	75.0	768	472	-
93.0	237	-	219	50.0	772	454	226
91.0	237	-	217	20.0	795	-	226
90.0	459	-	232	0.0	800	-	-

E : 216° tr.t. (1+1) 470°

a* : (1+5) II b* : (1+5) II - (1+5) I

t	p	t	p
95.0 %			
151	16.7	176	78.4
152	18.8	177	80.9
80.0 %			
175	1.0	286	94.7
223	12.1	314	203
247	17.9	340	350
50.0 %			
175	0.0	261	40.58
219	3.8	283	74.0
220	13.4	338	336

Sodium chloride (NaCl) + Niobium chloride ($NbCl_5$)

Morozov, Korchunov, 1956

%	tr.t. a*	tr.t. b*	tr.t. c*	%	tr.t. a*	tr.t. b*	tr.t. c*
100	207	-	-	79.0	-	259	426
97.0	200	-	-	76.0	-	258	431
96.0	207	234	-	70.0	-	258	430
94.0	206	256	-	58.0	-	253	440
93.0	206	254	-	40.0	-	253	-
89.0	207	256	-	30.0	-	248	-
85.0	206	256	-	20.0	-	236	-
82.0	-	256	432	0.0	-	-	800

a* : $NbCl_5$ + (1+1) - L + (1+1)

b* : (1+1) c* : NaCl + (1+1) - L + (1+1)

t	p_2	t	p_2
92 %			
156	18.2	178	54.4
158	23.4	252	166.0
69.7 %			
156	1.0	178	4.5
176	4.4	207	28.0
63.5 %			
157.5	1.7	212	37.2
175	3.9	245	156.5
176	4.2		
25 %			
159.5	2.1	220	46.1
181	6.6	246	116.8
207.5	29.8	272	295.0

Potassium chloride (KCl) + Rubidium chloride (RbCl)

Zhemchuzhni and Rambach, 1910

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	726	-	49.25	744	735
96.91	726.5	726	39.65	750	738
91.05	727	726	30.00	755	743
84.45	728	726	20.80	762	750
75.82	732	728	10.75	776	767
67.06	736	730	2.24	786.5	782
57.32	740	731	0	790	-

mol%	Q mix.
25	190
53	60
75	70

Keitel, 1925

%	f.t.	E	%	f.t.	E
0	776	-	60	715	715
25	745	728	65	716	715
35	732	720	75	716	715
50	718	725	100	722	-

Potassium chloride (KCl) + Cesium chloride (CsCl)

Zhemchuzhni and Rambach, 1910

mol%	f.t.	m.t.	tr.t.
100	646	-	451
96.83	642.5	640	446
94.0	639.5	635	441
88.54	633	623	437
81.92	625	618	-
74.89	620	618	-
66.04	616	616	-
51.73	625	618	-
40.07	655	637	-
29.69	690	666	-
16.03	729	712	-
6.87	769.5	752	-
3.41	780	770	-
0	790	-	-

Keitel, 1925

mol%	f.t.	E	tr.t.
100	629	-	449
85	616	611	430
65	611	610	-
20	707	690	-
0	776	-	-

Potassium chloride (KCl) + Thallium chloride (TlCl)

Sandonnini and Aureggi, 1911

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	776	-	70	535	-
10	742	-	75	525	-
20	712	-	80	490	429
40	654	-	90	470	430
50	616	-	100	429	-
60	582	-			

Poma and Gabbi, 1911

%	f.t.	E	min.	tr.t.	min.
100	415	-	-	-	-
97	395	142	130	-	-
95	375	141	170	-	-
88	310	142	520	-	-
75	181	143	920	-	-
72	144	144	1220	-	-
67	200	144	960	-	-
60	285	143	700	234	80
50	450	143	450	236	200
40	545	142	320	236	300
30	635	142	220	235	290
20	690	-	220	235	290
10	735	-	-	225	123
0	759	-	-	-	-

Korreng, 1914

%	f.t.	E	min.	tr.t.	min
100	425	-	-	-	-
95.99	400	143	280	-	-
91.40	360	150	420	-	-
86.19	326	151	570	-	-
79.94	260	150	660	-	-
76.45	219	150	720	-	-
72.65	162	150	770	-	-
69.49	191	152	760	-	-
63.91	244	150	700	-	-
57.61	370	150	600	244	90
53.24	430	148	560	243	120
47.96	514	150	450	240	140
40.91	586	147	180	238	160
22.78	690	146	-	338	70
0	775	-	-	-	-

Fontana, Gorin and al., 1952 (fig.)

mol%	f.t.	mol%	f.t.
100	425	54	245 (1+1)
80	295	50	350
65	150	40	500
60	200		

Potassium chloride (KCl) + Cuprous chloride (CuCl)

Sandonnini, 1911

mol%	f.t.	tr.t.	min.	E	min.
0	776	-	-	-	-
10	734	217	20	130	-
20	671	219	40	134	30
30	583	220	50	135	70
33.40	555	224	80	139	100
40	485	225	60	137	130
45	426	221	40	135	150
50	355	226	30	137	170
55	220	-	-	135	180
57.50	218	-	-	136	200
60	200	-	-	137	230
65	180	-	-	136	300
70	170	-	-	136	180
75	232	-	-	136	160
80	274	-	-	134	110
90	362	-	-	135	80
95	394	-	-	130	20
100	422	-	-	-	-

De Cesaris, 1911

%	f.t.	E	min.	tr.t.	min.
0	774	-	-	-	-
10	745	245	120	138	60
20	705	250	180	138	120
25	670	245	195	139	135
30	640	248	225	141	150
40	550	245	300	140	180
50	450	246	195	138	240
55	390	246	90	144	285
60	300	242	60	138	360
65	225	-	-	142	450
70	165	-	-	140	600
75	190	-	-	140	570
80	260	-	-	140	375
90	250	-	-	140	150
95	390	-	-	138	75
100	420	-	-	-	-

Potassium chloride (KCl) + Silver chloride (AgCl)

Zhemchuzhni , 1908

mol%	f.t.	E	mol%	f.t.	E
100	451	-	52.49	474	306
97.43	439	306°	49.88	494	"
93.32	426	"	47.51	512	"
88.67	405	"	45.12	535	"
84.06	384	"	40.00	562	"
79.25	353	"	35.27	596	"
74.30	329	"	27.39	638	"
74.30	330	"	22.10	667	"
74.8	319	"	17.70	700	"
68.73	383	"	11.03	730	"
62.73	444	"	5.48	760.5	305
55.00	457	"	0	790	-

Benrath and Wainoff, 1911

%	f.t.	%	f.t.
100	450	40	520
90	405	30	600
80	360	20	650
70	306	10	730
60	390	0	790
50	460		

Sackur, 1913

mol in 1000gr KCl	f.t.	mol in 1000gr KCl	f.t.
0	772	0.53	756.0
0.14	768.2	.55	755.6
.145	767.9	.575	754.4
.175	767.3	.63	754.3
.30	763.5	.71	754.1
.34	763.0	.76	748.7
.38	761.1	.95	748.4
.40	760.1	1.02	742.8
.435	757.1		

Tubandt and Abramowitsch (fig.)

mol%	f.t.	mol%	f.t.
100	440	40	560
90	400	30	620
80	370	20	690
70	335	10	760
60	390	0	810
50	485		

Lifshits, 1956.

%	f.t.
100	452
70	318 E
0	774

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ 600°	σ 700°	mol%	σ 600°	σ 700°
0	110	102.5	60	124.5	117
20	112.5	105	80	135	130
40	118	110	100	171.5	166.5

Harrap and Heymann, 1955

mol%	η 400°	η 450°
100	-	2330
80.6	-	2480
79.1	-	-
67.9	3130	250
57.7	-	-

mol%	η 500°	η 550°
100	2080	1850
80.6	2120	1840
79.1	-	-
67.9	2100	1800
57.7	-	-
46.8	-	-
44.9	-	1860

mol%	η 600°	η 650°
100	1660	1510
80.6	1620	1440
79.1	-	-
67.9	1560	1380
57.7	-	-
46.8	-	-
44.9	1610	1410
38.1	1630	1440
34.3	-	-

mol%	η 675°	η 700°
100	1450	1400
80.6	1360	1280
79.1	-	-
67.9	1300	1240
57.7	-	-
46.8	-	-
44.9	1320	1240
38.1	1350	1280
34.3	-	-

Next Page

Benrath and Wainoff, 1911

t	$\varkappa \cdot 10^6$	t	$\varkappa \cdot 10^6$
	0%		
640	5.0	720	36.9
650	6.5	730	43.7
660	8.5	740	59.6
670	12.0	750	77.9
680	13.4	760	91.9
690	18.6	770	133.8
700	20.9	780	199.2
710	28.8	790	1908000

t	10%	20%	30%
210	0.5	1.5	3.0
220	0.8	2.6	5.8
230	1.1	3.9	8.3
240	1.5	5.8	13.0
250	2.0	9.1	19.0
260	3.3	13.6	28.3
270	5.5	20.1	40.9
280	9.1	30.9	61.0
290	15.2	46.4	90.4
300	32.5	81.0	143.3
306	2513	6666	500000
t	**40%**	**50%**	**60%**
210	8.0	13.0	18.0
220	11.3	20.0	25.0
230	17.0	28.8	37.0
240	25.3	42.0	33.0
250	36.8	60.3	76.0
260	54.4	83.0	104.0
270	76.6	118.0	153.0
280	110.1	168.0	242.0
290	159.3	244.0	366.0
300	246.5	360.0	610.0
306	47500	75000	110000
t	**70%**	**80%**	**90%**
210	31.0	35.0	41.0
220	50.0	52.4	57.0
230	70.0	75.0	100.0
240	97.0	112.0	154.0
250	136.0	163.0	227.0
260	202.0	229.0	333.0
270	287.0	325.0	488.0
280	428.0	476.0	698.0
290	598.0	650.0	1036
300	800.0	990	1440
306	200000	200000	100000
t	**100%**	**t**	**100%**
170	5.4	320	2467
180	9.5	330	3247
190	17.5	340	4386
200	29.3	350	5905
210	45.0	360	8427
220	75.0	370	11029
230	121.0	380	18750
240	172.0	390	25861
250	271.0	400	33750
260	382.0	410	39473
270	539.0	420	50370
280	731.0	430	65200
290	1026	440	82600
300	1318	450	400000
310	1794	-	-

Tubandt and Abramovitch, 1927

mol%	$\varkappa$ 200°	250°	275°	300°
100	2.64	3.9	9.2	18.0
90	2.24	3.2	7.2	14.0
80	1.96	2.6	5.8	11.7
70	1.76	2.9	5.0	10.5
60	1.08	1.8	3.8	8.2
50	0.70	1.2	3.6	5.7
40	0.44	0.6	2.0	3.7
30	0.32	0.2	1.2	2.0
20	-	0.1	0.7	1.0
10	-	0	0.1	0.3

Harrap and Heymann, 1955

mol%	$\varkappa$	$\varkappa$
	400°	450°
100	-	37000
80.6	-	-
79.1	22980	25140
67.9	-	-
57.7	-	17340
	500°	550°
100	39020	40640
80.6	-	-
79.1	26960	28500
67.9	-	-
57.7	19300	20960
46.8	-	18440
44.9	-	-
	600°	650°
100	42120	43440
80.6	-	-
79.1	29960	31100
67.9	-	-
57.7	22440	23850
46.8	20060	21620
44.9	-	-
38.1	-	-
34.3	-	19860
	675°	700°
100	44040	44580
80.6	-	-
79.1	31720	32240
67.9	-	-
57.7	24480	25120
46.8	22380	23140
44.9	-	-
38.1	-	-
34.3	20760	26140

Stern, 1956.

t	Electrode Potentials (volts)	t	Electrode Potentials (volts)
	100 mol%		
476	0.9079	568	0.8767
507	0.8972	572	0.8755
511	0.8922	588	0.8711
534	0.8878	628	0.8556
	59.09 mol%		
568	0.9158	713	0.8820
600	0.9082	738	0.8760
681	0.8895	753	0.8724
	42.66 mol%		
666	0.9267	871	0.8782
716	0.9163	904	0.8678
788	0.8985	916	0.8645
810	0.8938		
	19.07 mol%		
704			0.9847
810			0.9737
914			0.9620
	9.821 mol%		
740			1.0540
844			1.0470
978			1.0242
	7.992 mol%		
778	1.0730	872	1.0626
816	1.0710	912	1.0626
862	1.0685		
	6.466 mol%		
829			1.1154
861			1.1127
906			1.1103

Potassium chloride (KCl) + Magnesium chloride ($MgCl_2$)

Menge, 1911 and 1912

%	f.t.	E	min.	
100	711	-	-	
93.7	684	468	7.3	
87.3	651	470	11.2	
77.6	609	472	15.5	
67.5	527	475	24.5	
63.1	-	473	30.3	
61.5	475	470	25.0	
59.6	479	473	15.8	
58.1	485	478	10.4	
52.2	484	432	5.0	
51.1	486	436	8.0	
47.3	464	433	20.3	
42.5	446	431	27.1	
39.0	437			(1+1)
37.5	449	423	31.1	
36.2	474	427	37.1	
33.5	512	426	31.0	
30.7	561	427	31.4	
25.2	616	426	23.0	
15.9	697	425	15.4	
9.3	740	426	9.3	
2.0	752	426	2.0	
0	776	-	-	

See also: Jänecke, 1912 and Menge, 1912

Klemm, Beyersdorfer and Oryschkewitsch, 1949 (fig.)

mol%	f.t.	E	mol%	f.t.	E
0	770	-	50	488	(1+1)
5	750	430	55	480	470
10	725	430	60	470	470
20	620	430	70	565	470
30	430	430	80	640	470
33.3	432	(2+1)	90	685	470
35	431	431	100	715	-
40	458	431			

Ivanov, 1952 (fig.)

%	f.t.	m.t.	tr.t.
100	710	-	-
90	665	455	-
80	610	468	-
70	535	467	-
62	467	467	-
57	490	E	425
50	480	440	425
41.2	440	436 tr.t.	-
40	438	436	429
39	437	435	430
38.5	436	435 E	-
37.4	435	435	425
35	490	435	425
30	560	435	425
20	655	435	425
10	720	435	
0	775		

Kordes, 1956

mol%	activity B	mol%	activity B
100	1.000	80	0.670
95	0.922	70	.486
90	.842	60	.296
85	.758		

Karpatchew and Stromberg, 1935

mol%	d					
	500°	550°	600°	650°	700°	750°
19.15	-	-	-	1.584	1.554	1.526
34.08	1.670	1.644	1.618	.594	.568	-
34.49	.674	.648	.620	.594	.654	-
35.84	.663	.656	.608	.580	-	-
46.97	.704	.672	.640	.610	1.576	-
55.58	.754	.718	.684	.650	.614	-
70.31	-	-	-	.700	.672	1.644

Karpachev, Stromberg and Pooltoratskaya, 1935

mol%	d 800°	750°	700°	650°
0	1.485	1.515	-	-
7.2	.508	.534	-	-
15.8	.509	.546	-	-
30.0	.523	.561	1.594	1.624
40.0	.516	:546	.575	.598
54.7	-	.608	.644	.719
60.0	1.506	.534	.562	-
87.4	.659	.688	-	-

Karpachev and Stromberg, 1935

mol %	η 500°	550°	600°	650°	700°	750°
11.19	-	-	-	-	-	92
27.42	-	-	-	157	125	110
26.05	-	179	144	115	99	-
34.49	269	211	170	140	-	-
35.53	278	208	151	128	-	-
35.65	235	187	155	140	-	-
37.88	202	177	149	120	-	-
45.50	-	191	155	142	122	-
52.26	-	201	157	177	121	-
56.60	326	259	219	200	-	-

Desyatnikov, 1956

mol %	σ 700°	800°	900°	dσ/dt
0	105	-	99	0.066
10	-	89.6	83	0.058
20	92	85.5	80	0.048
30	87	81.5	76	0.038
40	83	78.5	75	0.043
50	81	76.9	74	0.043
60	80	75.5	73	0.038
70	78	74.1	71	0.033
80	74	71.7	69	0.027
90	71	68.7	66	0.021
100	66	-	65	-

Karpachev and Stromberg, 1935

mol %	(in mhos) 850°	800°	750°	700°	650°
0.00	2.35	2.21	-	-	-
7.93	-	1.97	1.87	-	-
16.59	-	.80	.70	1.55	1.45
26.38	-	.55	.43	.34	.22
34.98	-	.42	.32	.23	.12
44.41	-	.41	.32	.23	.13
51.31	-	.38	.31	.21	.13
61.98	-		.31	.21	.11
72.30	-	-	.30	.21	.10
81.30	-	1.42	.32	.22	.11
92.43	-	.31	.21	.10	-
100.00	-	.18	.06	-	-

	600°	550°	525°	500°
26.38	1.09	-	-	-
34.98	.01	0.92	0.83	0.75
44.41	.01	.88	.83	-
51.31	.00	.88	.83	-
61.98	0.99	-	-	-

Markov, Delimarski and Panchenko, 1955

mol%	e	τ (.10_3)	mol%	e	τ (.10_3)
		718°			
100	2.531	-0.69	39.0	2.681	-0.555
62.2	.547	-0.66	29.8	.764	-0.52
49.6	.600	-0.658	28.8	.783	-0.49
40.4	.677	-0.56	21.1	.814	-0.45

Potassium chloride (KCl) + Calcium chloride ($CaCl_2$)

Menge, 1911

%	f.t.	E	min.	
100	777	-	-	
98	769	607	3.7	
95	750	631	6.7	
90	705	641	11.6	
85	700	639	15.5	
81	-	640	27.2	
80	655	641	23.7	
75	707	644	16.9	
70	732	633	7.6	
65	749	632	3.7	
62	753	632	2.4	
60	754	-	-	(1+1)
59.8	754	-	-	
58	752	590	3.9	
55	742	594	7.0	
50	725	596	14.2	
45	688	598	15.1	
40	658	600	19.2	
35	-	600	35.5	
30	626	600	30.1	
20	696	597	19.2	
10	748	598	10.0	
2	775	596	5.4	
0	776	-	-	

Moldenhauer and Andersen, 1913

%	f.t.	%	f.t.
100	780	50	690
85	630	40	630
80	670	30	640
60	740		
(1+1)			

Lantsberry and Page, 1920

%	f.t.	E	%	f.t.	E
100	745	-	67.5	725	575
90	670	608	65	723	580
82.5	615	-	60	713	580
80	655	615	50	670	583
77.5	695	615	40	590	-
75	725	600	30	640	590
72.5	720	600	15	710	590
70	725	-	0	760	-

Gromakov and Gromakova, 1955

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	775	775	60	715	630
20	640	583	70	645	610
25	590	583	72.5	-	610
25	583	-	75	630	611
30	645	583	80	664	610
40	713	611	90	728	652
50	739	739	100	775	-

Sandonnini, 1920

%	κ(in mhos)	%	κ(in mhos)
	800°		
100	2.006	55	1.480
90	1.772	50	.492
82	.620	40	.550
75	.554	30	.708
70	.501	20	.951
65	.477	0	2.301
60	.478		

%	d 800°	d 900°
100	2.057	2.009
90	1.977	1.927
82	.896	.850
75	.872	.822
70	.844	.788
65	.813	.756
60	.780	.734
55	.751	.694
50	.715	.667
40	.671	.613
30	.622	.562
20	.573	.517
0	.495	.434

Potassium chloride (KCl) + Strontium chloride ($SrCl_2$)

Vortisch, 1914

%	f.t.	E	min.	f.t. (1+2)	min.
0	775	-	-	-	-
19.11	727	594	70	-	-
34.71	671	595	150	-	-
44.65	608	596	260	-	-
47.68	-	596	280	-	-
50.02	597	(594)	(90)	-	-
51.53	597	-	-	-	-
53.38	596	568	20	-	-
58.64	586	574	130	-	-
63.50	580	576	250	-	-
68.02	604	575	190	-	-
72.21	620	575	100	-	-
76.13	631	575	60	-	-
79.79	637	570	20	-	-
80.96	-	-	-	638	200
80.23	661	(562)	-	638	160
86.45	696	(560)	-	637	130
89.48	741	-	-	636	90
95.03	807	-	-	628	40
100.00	870	-	-	-	-

Potassium chloride (KCl) + Barium chloride ($BaCl_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	790	33	686 (1+2)
10	755	40	670 (2+3)
20	710	45	700
24	680 (1+3)	50	730
30	685	100	960

Gemsky, 1913

%	f.t.	E	min.	tr.t.	min.
0	775	-	-	-	-
23.69	738	660	80	-	-
41.12	685	658	250	-	-
48.21	-	660	350	-	-
54.49	662	658	170	-	-
58.28	663	-	-	-	-
65.06	660	652	130	-	-
69.56	-	652	290	-	-
73.64	694	651	230	-	-
80.73	756	650	170	-	-
86.70	815	650	130	-	-
91.71	870	652	80	-	-
96.17	909	649	50	-	-
99.27	942	645	20	930	30
100	955	-	-	930	80

Potassium chloride (KCl) + Zinc chloride ($ZnCl_2$)

Vortisch, 1914

%	f.t.	E	min.	tr.t.	min.
0	775	-	-	-	-
33.02	708	656	150	-	-
48.21	-	657	250	-	-
58.28	660	-	-	-	-
65.06	653	646	120	-	-
77.35	721	646	160	-	-
86.70	805	645	110	-	-
99.27	933	638	20	923	20
100	955	-	-	922	80

Elchardus and Laffitte, 1932

mol%	f.t. cooling	E	f.t. heating	E
4.39	755	655	755	657
9.29	727	654	-	657
17.16	692	656	694	657
19.95	684	656	684	655
23.76	663	657	663	657
25.00	-	658	-	655
32.35	658	645	657	646
37.00	654	645	654	644
38.44	652	647	652	645
39.93	-	644	649	-
43.82	655	646	661	646
48.36	700	646	701	646
61.14	764	645	-	646

E: 42.75 mol% 645° (1+1) 656°

Gromakov and Gromakova, 1955

mol%	f.t.	E	mol%	f.t.	E
0	775	-	35	-	644
8.5	735	680	43	-	644
17	698	650	45	656	644
26	665	650	55	725	644
27	-	650	65	788	644
33.3	-	654	100	960	-

Peake and Bothwell, 1954

mol%	d	mol%	d
	800°		
0.0	1.500	33.6	2.217
8.4	.687	45.7	.419
13.1	.765	49.5	.480
24.4	2.050	63.3	.748
29.8	.123	79.1	.912
33.3	.184	100.0	3.12C

mol%	σ	mol%	σ
	800°		
0.0	96.2	37.3	114.3
3.9	98.4	42.3	117.0
8.4	101.7	47.7	121.5
13.5	102.8	62.3	133.0[a]
24.4	108.9	82.1	144.2[a]
29.4	111.0	100.0	162.6[a]
34.0	112.8		

Nikonova, Pavlenko and Bergman, 1941

mol%	f.t.	solid phase
100	320	A
95	316	"
90	308	"
85	301.5	"
80	293	"
75	280.5	"
71	262	E
70	268	(1+2)
65	273	"
60	262.5	"
55	238	"
54	228	E
52	239	(3+2)
50	243	"
48	248	"
46.5	250	tr.t.
45	310	(2+1)
43	350	"
40	395	"
35	442	"
32	439	"
31.5	432	E
30	450	B
29	469	"
25	535	"
20	626	"
15	694	"
0	775	"

Ugai and Chatilo, 1949 (fig.)

mol%	f.t.		
100	315		
80	290		
72	262 E		
66	274	(1+2)	congruent
53	230 E		
48	250	(3+2)	incongruent
40	390		
33	446	(2+1)	congruent
30	430 E		
25	530		

Bues, 1955

Raman spectra in melt.

Potassium chloride (KCl) + Cadmium chloride ($CdCl_2$)

Brand, 1911

%	f.t.	E	min.	tr.t.	min.
100	562	-	-	-	-
97.91	552	375	60	-	-
90.77	493	383	175	-	-
85.16	418	386	265	-	-
78.67	411	383	210	-	-
75.23	429	380	30	-	-
71.09	431	-	-	-	-
67.80	425	387	90	-	-
62.11	400	391	300	-	-
56.97	404	390	330	-	-
51.31	477	391	200	460	70
45.04	560	392	170	464	150
38.07	625	389	110	461	200
30.26	680	389	90	461	160
21.46	719	382	20	459	120
11.46	749	-	-	458	80
0	774	-	-	-	-

(1+1) (4+1)

Dergunov, 1949 (fig.)

mol%	f.t.
100	562
80	490
67	390 E
60	420
50	430 (1+1)
38	390 E
32	450 tr.t. (2+1)
30	500
20	630
10	718
0	775

Ilyasov, Bostandzhiyan and Bergman, 1956.

mol%	f.t.	mol%	f.t.
100	568	55.0	395
94	525	55.0	383 E
89	483	53.0	389 tr.t.
87.5	460	54.0	389
85.0	446	52.5	422
82.5	422	52.0	430
80.0	396 E	50.0	451
77.5	410	49.0	456
72.5	424	47.5	476
67.5	428	46.0	496
63.0	422 (1+2)	40.0	588
60.0	413		
57.5	404		

Bergman and Bakumskaya, 1956.

mol%	f.t.	mol%	f.t.
100	568	38	388 E
90.5	540	37.5	390
81.9	504	37.0	391
74	452	36.6	392 tr.t.
66.7	383 E	36.1	398
60	414	34.3	428
53.9	416	32.5	448
50.4	430(1+1)	31.6	462tr t
45.5	424	29.1	504 (4+1)
40.4	404	25	570

Lorenz, Frei and Jabs, 1908

t	d	
	66.7 mol%	33.3 mol%
600	2.715	2.181
650	.681	.147
700	.647	.112
750	.613	.078
800	.579	.044

0% : d= 1.963 - 0.00057 t

100% : d= 3.731 - 0.000685 t

Lorenz and Adler, 1928

t	d	t	d
20 mol%		40 mol%	
702	1.877	598	2.308
685	.885	609	.299
649	.903	705	.227
636	.953	720	.216
60 mol%		649	.273
612	2.676	651	.270
609	.677	80 mol%	
608	.679	600	3.17
690	.604	707	2.906
695	.595	703	.907
700	.593	700	.907
701	.591	650	.979
656	.633		
653	.635		
651	.636		

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ 600°	σ 700°	mol%	σ 600°	σ 700°
0	110.5	103	70	90	84.5
20	99.5	94.5	80	88.5	84.5
40	93	86.5	100	83	80.5
60	90	84.5			

Sandonnini, 1920

%	κ	%	κ
	800°		
100	22500	55	17850
90	21100	50	17330
80	19290	35	18520
75	18410	30	19110
70	17710	20	20410
65	17030	10	21630
60	16620	0	23010

Bues, 1955

Raman spectra in melt.

Potassium chloride (KCl) + Mercuric chloride ($HgCl_2$)

Belaiev and Mironov, 1952

mol%	f.t.	mol%	f.t.
100	283.0	64.0	198.0
95	276.5	62.5	201.0
90	271.5	60	213.0
87.8	268.0	57.5	222.0
85.5	263.0	56	229.0
83.3	259.0	54	234.0
81.2	254.5	51.5	237.0
79.3	249.0	50	239.0
77.4	244.0	47.5	242.0
75.6	238.0	45.5	244.0
74.4	232.0	42.5	245.0
70	203.0	39.9	318.0
68.5	189.0	37.5	371.0
67	193.0	37.5	448.0
65.5	197.0	34.3	523.0

E: 68% 182°

tr.t.(1+2) 63% 199° (2+1) 50% 239°

(3+2) 45% 245°

Potassium chloride (KCl) + Cupric chloride ($CuCl_2$)

Glocker and Dietmeier, 1938

Absorption and reflexion in crystals.

Potassium chloride (KCl) + Stannous chloride ($SnCl_2$)

Rack, 1913

%	f.t.	E	min.
100	239	-	-
99.154	238	192	40
97.92	234	194	160
95.82	223	201	240
93.53	206	201	410
91.06	204	201	200
88.43	208 (1+3)	-	-
87.04	207	175	140
85.60	205	179	180
83.59	203	180	-
82.55	200	181	400
80.94	180	180	560
79.26	187	179	340
77.51	192	179	300
75.71	208	177	150
71.81	224 (1+1)	-	(520)
69.74	310	224	500
67.57	385	224	480
62.94	481	224	440
57.84	541	224	400
52.19	580	223	300
38.91	679	222	240
22.06	752	221	120
0	777	-	-

Potassium chloride (KCl) + Lead chloride ($PbCl_2$)

Lorenz and Ruckstuhl, 1906

%	f.t.	min.	solid phase	E	min.
100.00	493	385	A	-	-
98.00	469	-	"	410	62
96.00	442	-	"	410	166
95.00	430	-	"	411	236
93.71	-	-	"	411	456
91.78	416	-	mixed crys-	411	258
90.00	427	-	tals	411	79
88.16	430	410	(1+1)	-	-
87.00	428	-	"	406	38
85.00	420	-	"	407	139
83.00	412	-	"	405	250
80.00	405	-	"	405	451
78.83	-	-	"	405	550
78.00	405	-	(2+1)	405	497
75.00	421	-	"	402	373
71.28	440	-	"	403	212
70.00	452	-	(4+1)	405	159
67.50	467	-	"	479	62
65.06	479	-	"	480	70
60.00	558	-	B	481	139
55.38	570	-	"	480	185
50.00	604	-	"	481	230
42.69	638	-	"	480	214
40.00	651	-	"	480	185
30.00	686	-	"	480	134
20.00	725	-	"	475	77
10.00	747	-	"		36
0:00	771	220	"	-	-

71.28 to 55.38% tr.t. = 440°

Treis, 1914

%	f.t.	E	min.	tr.t.	min.
100	496	-	-	-	-
97.10	475	428	80	-	-
93.72	440	429	340	-	-
89.61	437	429	250	-	-
88.18	440	-	-	-	-
87.38	437	411	20	-	-
84.83	434	411	130	-	-
82	420	411	320	-	-
78.86	430	411	370	-	-
75.31	460	413	260	-	-
71.31	483	412	150	-	-
66.75	538	410	80	490	250
65.09	550	410	50	490	300
61.51	578	408	30	490	260
55.42	620	-	-	491	240
48.24	677	-	-	490	220
39.09	710	-	-	490	200
29.29	730	-	-	488	160
0	775	-	-	-	-

(1+2) and (2+1)

Ugai and Chatilo, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	498	38	490 (2+1)
90	475	30	600
73	429 E	20	700
67	440 (1+2)	10	750
60	435	0	775
51.5	411 E		

Ilyasov, Bostandzhiyan and Bergman, 1956

mol %	f.t.	mol %	f.t.
100	496	48.2	426
90.5	475	45.5	439
81.9	439	42.9	454
77.8	421 E	40.4	472
74.0	428	38.0	482
70.2	434 (1+2)	36.5 (2+1)	490 tr.t
66.7	440	35.6	501
63.3	439	33.3	530
60.0	437	29.1	564
56.9	430	25.0	596
53.9	421	21.2	626
51.0	411 E		

Ilyasov and Bergman, 1956.

mol%	f.t.	mol%	f.t.
100	496	51	410 E
90.5	475	48.2	426
81.9	439	42.9	454
77.8	421 E	38	482
74	428	36.6	490tr.t(2+1
70.2	434	35.6	501
66.7	440 (1+2)	33.4	530
60	437	25	596
53.9	421	17.7	626
		0	772

Lorenz, Frei and Jabs, 1908

mol %	t	d	mol %	t	d
97	500	4.856	85	650	4.27
61	500	3.75	61	750	3.443
40	500	3.09	40	800	2.777
85	550	4.406	25	850	2.29
85	600	4.338	16.7	900	2.072

d = 2.3375 - 7/8000.t + (3.316 - 28.1/50.000t) N/100

Boardman, Palmer and Heyman, 1955 (fig.)

mol %	σ 500°	σ 600°
0	118	110
20	117	108.5
40	117	107
60	118.5	109
80	124	114
100	136	126

Markov, Delimarskii and Panchenko, 1954 (fig.)

mol %	e 550°	e 600°
100	1.258	1.215
80	.262	.231
70	.270	.235
60	.282	.252
50	.297	.270
40	.320	.295
35	.333	.309

Potassium chloride(KCl) + Manganese chloride ($MnCl_2$)

Sandonnini and Scarpa, 1913

mol%	f.t.	tr.t. I	min.	tr.t. II	min.
100	650	-	-	-	-
90	627	450	30	-	-
80	575	450	40	-	-
70	405	450	80	-	-
65	-	449	150	-	-
60	466	448	30	-	-
55	488	448	-	-	-
50	495	-	-	-	-
45	490	-	-	-	-
40	465	428	20	-	-
35	-	428	120	-	-
30	449	428	60	-	-
25	540	426	20	445	80
20	620	-	-	445	120
10	715	-	-	446	60
0	774	-	-	-	-

(1+1)

Nacvlichvili and Bergman, 1939

mol%	f.t.	mol%	f.t.
100	650	52	490
98	648	50	496
96	638	48	492
94	634	46	488
92	622	44	478
90	610	41	462
85	592	40	458
80	568	38	442
75	542	37	436
70	506	35	432
69	492	34	428 E
66	474	32.5	432
64	449 E	31	442
62	462	29	482
60	474	25	532
58	478	20	602
56	486	15	652
54	488 (1+1)	0	774

Potassium chloride (KCl) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	193.0	62.5	178.7
81	192.5 compl.	57.1	213.1
77	187.4	51.2	248.4
73	177.2	50.5	255.5
68.8	162.1	48.5	375.0
65.5	158.4 (1+1)		
L_1+L_2+C	99.6mol% - 81.8mol%		193.0°

Shvartsman, 1940

%	f.t.
67	128 E
(1+1)	257
-	257 E

Potassium chloride (KCl) + Antimonium chloride ($SbCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.	
100	73.4 I	76.2	57.0	
99.26	72.4	72.0	69.8	(1+1)
98.4	71.8	70.5	79.7	
95.6	69.8	67.4	94.8	
95.0	69.2	63.8	109.4	
94.5	68.9	62.1	113.3	
93.6	68.9	60.8	118.3	(3+2)
91.3	67.4 II	59.0	162.5	
88.2	64.6	53.2	250.5	
87.7	64.1	51.2	277.1	
85.0	66.2	45.3	320.0	
82.0	59.2	41.5	360.0	
81.0	60.3 III	38.3	433.5	
77.8	58.0	31.3	533.0	

Rubidium chloride (RbCl) + Cesium chloride (CsCl)

Keitel, 1925

mol%	f.t.	E	tr.t. II
0	722	-	-
30	675	664	-
65	646	635	-
87.5	637	630	436
89	635	630	439
100	629	-	443

Zhemchuzhni and Rambach, 1910

mol%	f.t.	m.t.	tr.t.
100	646	-	451
97	641	638	446
94.29	637	636	440
88.5	635	635	435
82.09	639	635	433
76.03	641	636	-
67.05	644	637	-
53.08	652	640	-
40.15	671	647	-
25.14	689	660	-
12.13	709	678	-
5.89	716.5	698	-
3.12	721	710	-
0	726	717	-

Rubidium chloride (RbCl) + Thallium chloride (TlCl)

Sandonnini and Aureggi, 1911

mol %	f.t.	m.t.
0	716	-
10	684	-
20	660	-
30	636	-
40	605	505
50	559	450
60	522	430
70	495	429
75	475	429
80	460	429
90	438	429
95	432	429
100	429	-

Rubidium chloride (RbCl) + Cuprous chloride (CuCl)

Sandonnini and Aureggi, 1912

mol%	f.t.	tr.t. I	min.	II	min.
0	716	-	-	-	-
5	694	232	-	-	-
10	668	237	-	-	-
20	605	243	50	192	-
30	527	255	80	183	50
33.3	485	250	80	190	80
35	470	250	80	186	60
40	412	260	40	192	90
45	300	250	60	185	110
50	274	248	20	190	170
53	220	-	-	180	120
55	210	-	-	183	160
58	200	-	-	180	250
60	180	-	-	180	250

mol%	f.t.	tr.t. III	min.	IV
0	716	-	-	-
5	694	-	-	-
10	668	-	-	-
20	605	-	-	106
30	527	-	-	105
33.3	485	148	-	105
35	470	-	-	104
40	412	150	-	105
45	300	-	-	104
50	274	-	-	-
53	220	150	-	-
55	210	-	-	-
58	200	-	-	-
60	180	148	-	-
63	172	147	80	-
65		150	100	-
68	E	149	250	-
70	180	150	380	-
75	210	148	250	-
80	278	150	200	-
90	370	150	160	-
100	422	145	100	-

Rubidium chloride (RbCl) + Silver chloride (AgCl)

Sandonnini and Aureggi, 1911

mol%	f.t.	E	min.
0	716	-	-
10	678	253	50
20	628	253	80
30	560	253	100
40	478	253	120
50	365	253	160
60	E	252	230
70	350	252	180
80	378	253	80
90	415	253	30
95	433	251	20
100	455	-	-

Rubidium chloride (RbCl) + Magnesium chloride ($MgCl_2$)

Klemm, Beyersdorfer and Oryschkewitsch, 1949 (fig)

mol%	f.t.	E	mol%	f.t.	E
100	710	-	35	470	470
90	675	-	33.3	480	(2+1)
80	625	-	30	475	446
70	540	509	24	446	446
67	509	509	20	545	446
60	530	509	10	690	446
50	542	(1+1)	0	715	-
40	520	470			

Markov and Panchenko, 1954 and 1955

mol %	f.t.	E	mol %	f.t.	E
100	711	-	33.1	476	-
78.7	625	508	31.7	477	-
74.2	598	509	30.4	470	460
68.3	548	510	29.6	464	458
64.5	512	-	28.0	463	-
61.6	524	510	27.1	468	-
58.1	536	509	25.9	473	459
54.4	546	511	24.6	484	458
49.5	549	-	23.7	491	458
43.8	543	-	22.7	527	464
40.1	525	-	19.4	576	463
37.5	511	472	17.5	595	458
36.2	486	473	0	718	-

E_1 : 65 mol% 510° E_2 : 35.5 mol% 472°

E_3 : 29 mol% 459° (1+1) 550° (2+1) 477°

Markov, Delimarski and Panchenko, 1955

mol%	e	$\tau(.10^3)$	mol%	e	$\tau(.10^3)$
		718°			
100	2.531	-0.69	41	2.705	-0.53
59	.577	-0.64	34.3	.782	-0.52
47.7	.656	-0.55			

e = electromotive force in volts.

Rubidium chloride (RbCl) + Calcium chloride ($CaCl_2$)

Gromakov, 1950

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	724	724	60	828	-
10	648	580	65	812	-
20	644	582	70	790	-
25	724	580	75	772	670
30	760	-	80	730	680
35	800	-	85	700	680
40	820	770	90	726	670
45	834	800	95	756	660
50	840	840 (1+1)	100	775	775
55	836	-			

Rubidium chloride (RbCl) + Strontium chloride ($SrCl_2$)

Gromakov, 1950

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	721	-	50	700	700
10	664	575	60	694	650
20	595	560	66.7	667	650
25	575	560	68	650	-
30	620	559	70	675	650
33.3	646	558	75	720	-
40	678	595	100	870	870
45	692	675			
(1+1)					

Rubidium chloride (RbCl) + Barium chloride ($BaCl_2$)

Gromakov, 1950

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	725	725	35	662	-
5	706	624	40	660	650
10	690	624	50	720	650
15	670	626	60	790	-
20	640	624	70	846	650
25	640	624	80	890	720
30	660	640	100	960	960
(33.3)	663	- (1+2)			

Rubidium chloride (RbCl) + Lead chloride ($PbCl_2$)

Treis, 1914

mol%	f.t.	E	min.	tr.t.	min.
100	496	-	-	-	-
97.76	483	406	75	-	-
92.87	446	410	230	-	-
89.07	416	410	390	-	-
84.29	420	410	390	-	-
80.35	420	407	390	312	40
77.52	414	407	210	325	50
73.76	434	406	95	320	140
69.69	440	-	-	320	185
65.29	432	414	160	320	90
60.52	420	414	280	320	110
53.48	469	410	50	447	245
48.43	514	-	-	449	220
43.39	555	-	-	448	190
20.35	685	-	-	448	110
0	724	-	-	-	-

Markov, Delimarski and Panchenko, 1954 (fig.)

mol%	e 550°	e 600°
100	1.258	1.215
80	.265	.235
70	.275	.345
60	.285	.26
50	.305	.29
41	.34	.31
35	.35	-

Rubidium chloride (RbCl) + Cadmium chloride ($CdCl_2$)

Dergunov, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	562	33	440
80	480	31	436 E
73	400 E	20	595
60	475	10	690
50	500 (1+1)	0	718
47	434 (2+1)		

Gromakov, 1950

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	726	726	50	466	466
10	616	472	55	458	422
15	578	471	60	440	368
20	532	472	65	416	368
25	472	471 (1+3)	70	380	368
30	456	412	75	400	370
35	428	410	80	436	368
40	426	408	100	568	568
45	452	410			

Rubidium chloride (RbCl) + Manganese chloride ($MnCl_2$)

Nasvlichvili and Bergman, 1939

mol %	f.t.	mol %	f.t.	mol %	f.t.
100	650	50	552	31	458
95	610	45	544	30	446
90.2	584	40	498	29	436 E
80.6	544	39	494	27.5	466
75	510	36	474	27	486
70	465	35	466	25	330
69	460 E	34	458	20	592
65	494	33	454 E	15	634
60	532	32	466	0	717
55	544				

Gromakov, 1950

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	726	726	40	498	440
10	650	474	45	520	-
15	606	476	50	526	(1+1)
20	568	476	55	524	-
22	540	-	60	510	460
25	506	476(1+3)	65	482	462
27	476	-	70	470	458
30	468	437	75	500	460
32	460	440	80	530	456
35	440	-	100	650	650

Cesium chloride (CsCl) + Thallium chloride (TlCl)

Sandonnini and Scarpa, 1912

mol%	f.t.	tr.t. I	tr.t. II	m.t.
0.0	639	-	450	-
5.0	619	-	452	-
10	600	-	459	-
15	572	466	-	-
20	563	466	-	-
25	542	464	-	-
30	532	465	-	-
35	520	466	-	-
40	501	466	-	-
50	450	-	-	390
60	408	-	-	390
65	400	-	-	390
70	393	-	-	390
75	390	-	-	-
80	400	-	-	390
90	417	-	-	400
100	429	-	-	-

Cesium chloride (CsCl) + Cuprous chloride (CuCl)

Sandonnini and Scarpa, 1912

mol%	f.t.	E	min.	tr.t.I	min.	tr.t.II
0.0	639	-	-	-	-	450
5.0	618	-	-	320		452
10	595	-	-	322	20	449
20	541	-	-	320	30	-
30	428	-	-	318	60	-
35	372	-	-	320	80	-
40	348	-	-	318	120	-
45	298	234	30	-	-	-
50	276	236	70	-	-	-
55	-	235	180	-	-	-
60	260	235	40	-	-	-
65	269		-	-	-	-
66.6	264	-	-	-	-	-
70	268	218		-	-	-
75	248	215	50	-	-	-
80	268	216	80	-	-	-
90	350	214	60	-	-	-
100	422	-	-	-	-	-

Cesium chloride (CsCl) + Silver chloride (AgCl)

Sandonnini and Scarpa, 1912

mol %	f.t.	E	min.	tr.t.I	min.	tr.t.II
0.0	639	-	-	-	-	450
10.0	602	-	-	305	50	449
20.0	550	-	-	310	70	450
30.0	473	-	-	310	100	450
33.3	430	-	-	310	150	-
40.0	396	-	-	310	150	-
50.0	320	-	-	310	160	-
55.0	307	256	20	-	-	-
60.0	298	260	60	-	-	-
65.0	273	258	100	-	-	-
70.0	266	256	130	-	-	-
80.0	372	258	200	-	-	-
90.0	412	249	80	-	-	-
100.0	455	-	-	-	-	-

Cesium chloride (CsCl) + Magnesium chloride ($MgCl_2$)

Markov and Panchenko, 1955

mol%	f.t.	f.t. (1+3)	E
100	711	-	-
79.4	614	567	-
76.6	589	565	-
74.4	578	564	-
72.4	558	-	-
71.9	551	-	-
70.8	554	-	546
69.4	548	-	548
64.3	575	-	548
59.5	594	-	548
54.6	606	-	-
49.9	610	-	-
43.0	598	-	543
39.2	573	-	542
33.4	545	-	-
30.0	542	-	518
28.5	528	-	-
26.7	523	-	-
24.4	527	-	-
23.6	523	-	-
20.3	518	-	510
18.0	535	-	512
15.2	552	-	512
0	643	-	-

(1+1) 610° (2+1) 545°

(3+1) 527° (1+3)

E_1 : 69.4 mol% 548°

E_2 : 37 mol% 542°

E_3 : 26.5 mol% 518°

E_4 : 21.5 mol% 511°

Cesium chloride (CsCl) + Cadmium chloride ($CdCl_2$)

Dergunov, 1949 (fig.)

mol%	f.t.
100	562
80	490
75	440 E_1
60	525
50	545 max(1+1)
40	505
35	454 E_2
32	462 max(2+1)
20	445 E_3
10	595
0	630

Cesium chloride (CsCl)+ Lead chloride ($PbCl_2$)

Gromakov, 1950

mol%	f.t.	mol%	f.t.
0	645	50	590 (1+1)
10	588	54	582
14	559	58	567
16	540	60	556
18	522	62	548
20	504	64	538
22	478	66	527
24	468	68	515
26	472	70	504
28	472	72	486
30	485	74	468
32	506	76	445
34	522	78	430
36	533	80	413
38	544	82	398
40	559	84	410
42	568	86	421
44	573	88	432
46	579	90	440
48	584	95	470
		100	500

Thallium chloride (TlCl) + Cuprous chloride (CuCl)

Sandonnini, 1911

mol%	f.t.	tr.t.	min.	E	min.
0.00	429	-	-	-	-
5.00	410	-	-	-	-
10.00	380	223	-	-	-
15.00	370	223	-	-	-
21.15	345	222	40	-	-
30.00	300	226	140	-	-
33.40	283	226	180	-	-
37.75	270	224	130	-	-
40.00	260	223	100	123	20
45.00		221	60	124	60
50.00	218	-	-	121	80
50.10	208	-	-	123	90
55.00	173	-	-	124	110
61.70	166	-	-	122	240
64.50	166	-	-	123	200
65.00	183	-	-	123	190
70.10	253	-	-	120	130
81.30	313	-	-	120	100
90.06	362	-	-	116	30
95.00	385	-	-	116	-
100.00	422	-	-	-	-

Thallium chloride (TlCl) + Silver chloride (AgCl)

Sandonnini and Aureggi, 1911

mol %	f.t.	tr.t.I	min.	tr.t.II	min.
100	455	-	-	-	-
95	430	-	-	210	-
90	416	-	-	210	30
75	340	-	-	210	70
60	221	-	-	210	150
55	222	-	-	210	80
50	235	-	-	208	40
45	249	-	-	208	20
40	253	253	140	209	-
35	284	252	100	-	-
33.5	297	252	90	-	-
30	321	252	80	-	-
25	345	250	40	-	-
10	402	249	20	-	-
0	429	-	-	-	-

Sandonnini, 1913 and 1920

%	ϰ (melt)(in mhos)	%	$ϰ.10^6$ (solid)
	500°		200°
100	3.653	100	50
77.23	2.925	77.23	65
58.00	2.260	58.00	86
37.00	1.771	47.30	104
20.57	1.470	37.00	92
0	1.215	28.52	75
		20.51	65
		0	13

Thallium chloride (TlCl) + Beryllium chloride ($BeCl_2$)

Schmidt, 1929 (fig.)

mol%	f.t.	mol%	f.t.
0	432	55	320
10	400	60	340
18	370	66	365
20	400	80	355
30	455	85	350
33	458	90	370
40	445	100	404
50	365		

Thallium chloride (TlCl) + Magnesium chloride ($MgCl_2$)

Korreng, 1914

%	f.t.	E	min.	tr.t.	min.
0	435	-	-	-	-
2.05	425	360	120	-	-
6.56	409	361	240	-	-
16.59	412	362	310	-	-
28.46	495	357	180	-	-
44.30	523	-	-	498	290
54.40	585	-	-	499	240
78.16	677	-	-	499	140
100.00	718	-	-	-	-

(1+1)

Thallium chloride (TlCl) + Calcium chloride ($CaCl_2$)

Korreng, 1914

%	f.t.	E	min.
0	435	-	-
2.38	425	419	350
7.56	566	419	400
19.98	674	417	270
31.67	683	-	-
41.02	677	645	120
51.96	665	648	200
69.49	-	647	175
72.43	737	647	160
100	782	-	-

(1+1)

Thallium chloride (TlCl) + Strontium chloride ($SrCl_2$)

Korreng, 1914

%	f.t.	E	min.	tr.t.	min.
0	435	-	-	-	-
3.37	428	417	360	-	-
6.85	421	414	540	-	-
14.20	471	415	500	-	-
22.10	527	416	450	-	-
26.28	553	416	400	-	-
30.62	560	416	300	-	-
39.84	585	416	220	569	100
49.83	667	416	120	568	100
60.71	723	415	100	566	70
78,96	793	-	-	563	50
100	872	-	-	-	-

(1+1)

Thallium chloride (TlCl) + Barium chloride ($BaCl_2$)

Korreng, 1914

%	f.t.	tr.t.	E	min.
0	435	-	-	-
6.59	563	-	429	600
13.31	642	-	430	530
22.48	687	-	431	500
36.70	-	-	430	410
48.96	843	-	430	300
66.99	-	-	426	190
78.76	942	929	429	150
92.04	-	-	426	50
97.71	953	925	423	-
100	962	925	-	-

(2+1)

Thallium chloride (TlCl) + Zinc chloride ($ZnCl_2$)

Korreng, 1914

%	f.t.	E	min.	tr.t.
0	435	-	-	-
5.95	403	332	180	-
10.77	370	334	280	-
15.95	339	336	340	-
19.61	343	334	200	-
22.15	352	-	-	-
27.51	332	193	170	170
36.27	234	191	220	174
44.01	216	193	150	176
53.23	226	-	--	162
63.07	244	214	260	-
69.48	259	215	180	161
76.33	269	214	150	156
88.32	272	211	40	158
100	275	-	-	-

(2+1)

Sementsova, Bergman and Lesnykh, 1956

mol%	f.t.	mol%	f.t.
0	430	42.9	426
5.9	420	48.2	430
11.2	400	53.9	430
17.7	364	60.0	420
21.2	336	66.7	410
23.1	326	74.0	474
25.0	340	90.5	546
33.3	392	100.0	567

(1+1)

Sandonnini, 1920

%	κ (in mhos)	%	κ (in mhos)
	600°		
100	1.971	40	1.520
80	1.808	25	1.564
65	1.781	15	1.610
60	1.665	10	1.664
50	1.566	0	1.702
46.3	1.522		

Thallium chloride (TlCl) + Cadmium chloride ($CdCl_2$)

Korreng, 1914

%	f.t.	E	min.
0	435	-	-
7.84	400	315	330
16.06	325	314	600
24.70	374	315	520
33.79	423	315	270
43.36	436	-	-
53.45	422	407	450
64.11	449	408	600
75.38	510	408	420
87.33	553	408	220
96.76	573	407	100
100	578	-	-

(1+1)

Sandonnini, 1913

mol%	f.t.	E	min.
0.0	429	-	-
10.0	383	299	20
20.0	331	298	50
25.0	315	299	80
30.0	316	299	70
36.5	372	298	40
47.0	421		-
50.0	426	-	-
57.5	424	400	50
67.5		400	110
80.0	481	400	60
85.0	504	401	40
90.0	525	401	30
95.0	550	400	-
100.0	568	-	-

(1+1)

Thallium chloride (TlCl) + Mercuric chloride ($HgCl_2$)

Sandonnini, 1913

mol%	f.t.	m.t.	E	min.	tr.t.	min.
0	429	-	-	-	-	-
3.3	402		-	-	-	-
6.2	375	-	-	-	250	30
15.9	338	-	-	-	250	60
21.0	275	-	-	-	250	40
25.9	243	-	-	-	-	-
30.2	224	-	201	-	-	-
35.0	213	-	200	70	-	-
38.2	220	-	203	30	-	-
43.2	222	-	-	-	-	-
48.8	224	-	183	-	-	-
55.0	217	-	183	20	-	-
60.8	200	-	183	100	-	-
66.0	190	-	-	60	-	-
76.4	230	190	-	-	-	-
87.5	255	235	-	-	-	-
92.7	270	250	-	-	-	-
100.0	275		-	-	-	-

(1+1)

Thallium chloride (TlCl) + Stannous chloride ($SnCl_2$)

Sandonnini, 1913

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0.0	429	-	60.0	230	-
10.0	401	-	70.0	210	200
20.0	360	-	80.0	190	-
30.0	320	300	90.0	212	190
40.0	194	260	100.0	250	-
50.0	265	252			

(1+1)

Korreng, 1914

%	f.t.	E	min.
0	435	-	-
6.04	403	299	110
12.28	351	299	240
20.91	310	-	-
28.40	293	235	170
39.35	242	233	240
44.23	244	-	-
56.93	222	179	230
76.03	206	178	270
93.78	235	178	50
100	241	-	-

Thallium chloride (TlCl) + Lead chloride ($PbCl_2$)

Sandonnini, 1913

mol%	f.t.	E	mol%	f.t.	E
0.0	429	-	60.0	420	-
10.0	409	386	70.0	436	-
20.0	389	376	80.0	460	420
30.0	380	-	90.0	480	450
40.0	373	-	100.0	495	-
50.0	388	-			

(1+1)

Korreng, 1914

%	f.t.	E	min.
0	435	-	-
11.43	406	388	360
22.50	402	387	240
27.90	407	-	-
33.23	403	377	40
43.63	385	378	380
53.73	407	378	270
63.52	432	377	100
69.90	435	-	-
73.04	434	427	140
82.28	447	428	420
91.27	479	426	200
100	500	-	

(1+1)

Favorskii, 1941

mol %	f.t.	mol %	f.t.
100	498	46.0	386
95.0	487	45.1	383
90.2	476	44.1	379
85.2	456	43.0	375
83.2	452	42.0	371
81.4	446	41.4	375
79.5	437	40.0	376
77.0	427	38.0	382
75.2	421	35.0	386
73.4	422	34.3	389
71.0	426	31.0	392
69.2	427	29.1	395
66.7	427	27.0	396
65.1	426	25.0	397
62.2	425	23.0	396
60.0	422	20.9	395
58.2	420	19.0	392
56.1	415	17.0	386
54.0	412	14.9	380
52.1	408	13.5	382
50.1	402	9.8	397
49.1	397	5.1	415
48.2	397	0	428
47.0	392		

Thallium chloride (TlCl) + Manganese chloride ($MnCl_2$)

Nacvlichvili and Bergman, 1939

mol %	f.t.	mol %	f.t.
100	650	41	475
93	626	39	471
85	588	37	461
78	562	35	449
76.5	543	34	440
73	534	30.5	417
71.8	523	28	395
69	513	25	366
67.5	502	22.5	347
66	491	21.5	328 E
64	480	21	336
62	460 E	19.5	348
61	471	17.7	361
60	480	16	373
58	486	13.5	385
56	489	12	391
54	492	10.5	398
52	496	9	401
50	497	8	407
48.8	495	7	410
47	493	6	414
45	490	5	418
43	485	0	427

(1+1)

Gromakov, 1951

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	426	426	55	517	-
10	395	340	60	500	492
20	340	323	(64)	492	-
(23)	320	-	65	500	492
25	380	323	70	538	-
30	465	-	80	590	-
40	513	348	100	650	650
50	524	-			

Thallium chloride (TlCl) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.		mol%	f.t.	
99.48	189.4		66.9	196.6	(1+1)
99.10	189.9	compl.	65.8	211.4	
84.4	190.2		62.1	248.1	
81.3	182.3		57.9	274.1	
78.2	172.8		53.8	290.0	
73.9	158.8		48.8	295.5	
72.8	160.5	(1+2)	46.5	288.1	
70.8	162.9		44.5	325.0	
68.6	172.6		43.5	360.0	

98.8 - 85.3mol% $C + L_1 + L_2 =$ 192.0°

Thallium chloride (TlCl) + Bismuth chloride ($BiCl_3$)

Scarpa, 1912

mol%	f.t.	tr.t. I	min.	tr.t. II	min.
0	429	-	-	-	-
5.0	410	360	30	-	-
10.0	382	360	70	-	-
12.5	-	360	100	-	-
15.0	370	361	70	-	-
17.5	382	360	50	-	-
20.0	396	360	30	-	-
25.0	413	-	-	-	-
30.0	404	331	-	-	-
35.0	380	326	-	-	-
37.5	370	330	-	223	30
40.0	342	300	-	221	40
45.0	293	-	-	224	60
50.0	196	-	30	225	40
55.0	176	-	50	225	40
60.0	155	-	100	-	-
65.0	-	-	150	-	-
67.5	162-165	-	180	-	mixed crys
70.0	168-155	-	20	-	"
75.0	178-160	-	-	-	"
80.0	187-167	-	-	-	"
85.0	204-184	-	-	-	"
90.0	217-195	-	-	-	"
95.0	217-195	-	-		"
100.0	224	-	-	-	"

(3+1) (2+1) (3+2)

Thallium chloride (TlCl) + Ferric chloride ($FeCl_3$)

Scarpa, 1912

mol%	f.t.	tr.t.	min.
0.0	429	-	-
10.0	385	266	(30")
19.0	355	260	(60")
22.0	325	267	(120")
26.0	-	262	(160")
29.0	276	261	(70")
33.0	290	-	-
35.0	252	221	(40")
37.0	233	217	(50")
45.0	260	218	(40")
52.0	280	220	(30")
59.0	290	221	(30")
62.0	299	218	(20")
100.0	302	-	-

(2+1)

Cuprous chloride (CuCl) + Silver chloride (AgCl)

Poma and Gabbi, 1911

%	f.t.	E	min.
0	415	-	-
4	410	-	-
6.66	400	-	-
10	395	-	-
12.5	388	240	20
13.3	385	245	30
25	359	250	200
37.5	324	"	320
45	300	"	400
50	279	"	460
55	260	"	520
62.5	270	"	460
67	295	"	400
75	335	"	280
85	385	"	180
90	405	"	76
93.33	415	"	-
97	435	"	-
100	451	-	-

Sandonnini, 1911

mol%-	f.t.	E	min.
100	455	-	-
92.00	433	-	-
86.25	412	260	-
73.50	371	262	40
62.70	323	262	140
50.25	-	260	250
45.40	E	261	250
41.60	285	261	150
31.70	314	260	60
21.40	350	260	30
14.80	365	253	-
6.00	398-368	-	-
3.50	410-391	-	-
0	422	-	-

Urazov and Chelidze, 1941

%	f.t.	E	min.
100	455	-	-
95	445	-	-
90	414	-	-
85	389	253	1
83	386	252	4
70	327	250	8
67	311	252	10
55	-	252	14
50	267	253	12
45	298	252	10
40	308	252	8
27	367	252	4
13	389	252	-
0	422	-	-

Barth and Lunde, 1926

Lattice constants.

Cuprous chloride (CuCl) + Mercurous chloride (HgCl)

Jänecke, 1923

%	f.t.	%	b.t.
0	425	50	409
10	418	60	395
20	405	72	380
30	390	80	380
40	373	90	380
50	360	100	380
60	341		
65	330 E		
70	365		
72	380		

Cuprous chloride (CuCl) + Magnesium chloride ($MgCl_2$)

Menge, 1911

%	f.t.	E	min.
100	711	-	-
94.3	702	408	1
66.2	655	407	7.7
45.1	614	404	11.9
25.8	564	395	19.3
8.2	496	408	27.1
5.8	456	406	31.0
4.2	419	408	30.7
1.7	418		
0	418	-	-

Cuprous chloride (CuCl) + Calcium chloride ($CaCl_2$)

Menge, 1911

%	f.t.	E	min.
100	777	-	-
90	756	397	2.4
80	723	385	5
70	722	394	7.9
50	690	391	12.7
30	634	395	17.3
15	483	398	23.4
10	-	400	23.3
9	394	386	17.9
7	397	389	16.9
5	402	391	14
0	418	-	-

Cuprous chloride (CuCl) + Zinc chloride ($ZnCl_2$)

Herrmann, 1911

%	f.t.	m.t.	E	min.
0	424	-	-	-
10	410.6	348	-	-
20	395	-	241	1-
30	378	-	242	3.7
50	343	-	232	7
70	295	-	242	10
90	-	-	243	15
92.5	250	-	239	4
95	-	243	-	-
100	261.5	-	-	-

Palkin and Chepurko, 1956, fig.

mol%	f.t.	mol%	f.t.
0	434	60	340
10	420	70	320
20	400	80	300
30	395	88	283 E
40	390	94	325
50	380	100	319

Cuprous chloride (CuCl) + Stannous chloride ($SnCl_2$)

Herrmann, 1911

%	f.t.	E	min.
0	424	-	-
10	402	175	2
30	354.5	172	8
50	289.5	172	14
70	211	172	17
77.5	-	172	22
90	218	170.7	9
95	233	168	5
100	247.2	-	-

Cuprous chloride (CuCl) + Lead chloride ($PbCl_2$)

Herrmann, 1911

%	f.t.	E	min.
100	501	-	-
90	431.5	274.5	4.5
80	365	281	7.6
70	301	280	11.8
60	290	280	13
50	314.5	283.5	10.5
40	346	282	8.8
30	369	281	6.2
20	391.5	282.5	4.8
10	414.5	281	2.5
0	424	-	-

Pélabon and Laude, 1929

mol%	f.t.
100	498
33	258 E
0	440

Cuprous chloride (CuCl) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	291.5	52.1	230.9
89.5	190.0	51.9	231.6
71.0	187.4	49.1	231.6
69.2	175.3	46.0	231.9
67.8	178.3 (1+1)	42.6	225.9
59.4	209.9	39.9	252.5
		38.0	291.5

Cuprous chloride (CuCl) + Bismuth chloride ($BiCl_3$)

Herrmann, 1911

%	f.t.	E	min.
0	424	-	-
10	413	185.5	3.5
30	385.1	188	7
50	342	191.5	11
70	275	187.0	15
80	214	190	17
90	203.5	195.2	13
95	210.5	193.5	5
100	224	-	-

Cuprous chloride (CuCl) + Ferric chloride ($FeCl_3$)

Herrmann, 1911

%	f.t.	E	min.
0	424	-	-
10	395	307	6
30	357	307	14
50	-	304	6.4
60	318	306	4
62	320	-	-
70	315	263	4.5
80	305	262	13
90	273	263	15.7
95	284	263	8
100	298	-	-

Silver chloride (AgCl) + Mercurous chloride (HgCl)

Jänecke, 1923

E : 55 % 250°

Silver chloride (AgCl) + Beryllium chloride ($BeCl_2$)

Schmidt, 1929

mol%	f.t.	E	min.
0	461	-	-
10	432	218	3.5
20	480	218	10
30	317	233	15
40	235	235	22
50	293	235	12
60	330	233	8
70	338	227	6
80	361	227	4
90	381	216	1.5
100	404	-	-

E: 235° 40 mol%

Silver chloride (AgCl) + Magnesium chloride ($MgCl_2$)

Menge, 1911

%	f.t.	E	min.
100	711	-	-
94.4	701	446	2.7
66.6	659	452	10.0
45.4	602	452	16.9
21.8	522	450	18.8
17.2	509	452	21.2
13	482	452	16.7
8.7	-	452	25.0
6.1	454		
2.8	455		
0	455	-	-

Silver chloride (AgCl) + Calcium chloride ($CaCl_2$)

Menge, 1911

%	f.t.	E	min.
100	777	-	-
90	740-748	448	4.6
80	710-719	444	11.6
70	685-691	484	12.4
50	662	445	20.3
30	608	450	25.4
20	531	449	22.6
15	-	451	27.0
9	453	449	14.6
5	452	450	13
0	455	-	-

Silver chloride (AgCl) + Zinc chloride ($ZnCl_2$)

Glistenko and Artemova, 1953 (fig.)

mol%	f.t.	mol%	f.t.
0	456	42.9	248 E
11.2	425	66.7	295
25.0	375	100	318

Silver chloride (AgCl) + Cadmium chloride ($CdCl_2$)

Wagner and Hantelmann, 1950

mol%	f.t.	E	mol%	f.t.	E
0	445	445	50	450	435
10	455	-	60	465	435
20	445	-	80	510	435
30	440	435	100	565	565
40	440	435			

Silver chloride (AgCl) + Mercuric chloride ($HgCl_2$)

Bergman and Gönsk , 1926

mol%	f.t.	E	mol%	f.t.	E
100	281	-	33	381	272
90	276	272	23	404	"
82	274	"	14	425	"
67	305	"	5	442	"
54	338	"	0	451	"

Silver chloride (AgCl) + Lead chloride ($PbCl_2$)

Urazov and Karnaukhov, 1956

%	f.t.	t.	E
100	497	-	-
89.75	447	-	313
75.0	392	-	314
65.0	351	-	314
56.4	-	-	314
46.4	348	-	313
35.0	398	-	314
25.0	408	-	315
15.0	435	-	314
0	455	-	-

Treis, 1914

%	f.t.	E	min.
100	496	-	-
97.36	477	307	50
88.58	437	310	140
78.27	392	310	235
65.98	341	309	300
61.34	322	310	340
51.09	322	308	335
32.65	383	310	215
9.26	431	306	70
0	457	-	-

Tubandt and Eggert, 1920

mol%	f.t.	E	mol%	f.t.	E
0	455	-	60	391.5	310
10	421	313	70	419	311
20	384	312	80	447	310
30	347	312	90	473	311
40	-	314	100	498	-
50	351.5	311			

Pélabon and Laude, 1929

mol%	f.t.
100	498
39	291 E
0	455

Tarasenkov and Bogoslovskaya, 1935

%	p			
	500°	600°	700°	800°
1.0	-	-	-	0.141
22.7	-	-	2.84	-
23.0	-	0.317	-	-
24.3	-	-	-	12.82
48.5	0.665	-	-	-
49.0	-	0.877	-	-
49.5	-	-	8.74	-
68.5	-	-	-	11.55
70.5	-	1.44	-	-
73.0	0.680	-	-	-

Salström, 1934

t	d	t	d
50 mol%			
428.1	5.013	502.5	4.903
436.3	4.489	521.9	.878
452.8	.969	536.3	.860
462.7	.956	546.3	.848
476.5	.937	559.1	.831
487.9	.922	576.4	.810

Boardman, Dorman and Heymann, 1949

mol%	molar vol. in cc
600°	
100	57.92
80.6	52.60
70.9	49.93
61.1	47.28
53.4	45.10
42.6	42.15
20.3	36.04
0	30.50

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ		mol%	σ	
	500°	600°		500°	600°
0	175	171	60	143.5	134.5
20	160	152	80	139.5	129.5
40	150	141	100	136	126

Tubandt and Reinhold, 1923

mol%	κ	mol%	κ
	240°		
0	14.0	50	6.2
6	11.0	60	5.7
20	8.9	70	5.1
30	7.7	80	4.7
40	6.9	90	4.1

Harrap and Heymann, 1955

mol%	η	$\varkappa$	η	$\varkappa$
	350°		400°	
59.8	-	-	-	1.272
40.2	-	1.374	-	1.668
39.2	6450	-	4630	-
19.8	-	-	3540	-
13.8	-	-	-	2.620
	450°		500°	
100	-	-	4560	1.430
82.6	-	-	4130	-
80.3	-	1.318	-	1.594
62.0	4580	-	3470	-
59.8	-	1.560	-	1.816
40.2	-	1.946	-	2.196
39.2	3620	-	2950	-
19.8	2930	-	2470	-
13.8	-	2.882	-	3.112
0	2330	3.700	2080	3.902
	550°		600°	
100	3540	1.694	2750	1.920
82.6	3180	-	2590	-
80.3	-	1.840	-	2.072
62.0	2770	-	2300	-
59.8	-	2.050	-	2.266
40.2	-	2.406	-	2.606
39.2	2470	-	2120	-
19.8	2110	-	1840	-
13.8	-	3.298	-	3.464
0	1850	4.064	1660	4.212
	650°		700°	
100	2220	2.132	1870	2.334
82.6	2150	-	1990	-
80.3	-	2.296	-	2.508
62.0	1970	-	1840	-
59.8	-	2.480	-	2.680
40.2	-	2.790	-	2.956
39.2	1860	-	1760	-
19.8	1690	-	1650	-
13.8	-	3.600	-	3.732
0	1510	4.344	1400	4.458

Mulcahy and Heymann, 1943 (fig.)

mol%	λ	mol%	λ
	500°		
0	110	60	54
20	80	80	48
40	65	100	40

Silver chloride (AgCl) + Manganese chloride ($MnCl_2$)

Gromakov, 1951

mol%	f.t.	mol%	f.t.
0	455	35	466
5	454	40	484
10	452	45	502
15	451	50	520
20	450	60	552
25	448	70	585
30	445	100	650

Silver chloride (AgCl) + Aluminum chloride($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	190.2compl.	60.3	121.9
99.62	191.2	60.1	122.2 (1+1)
79.9	190.1	57.4	138.9
73.5	184.5	53.0	146.5
67.4	161.5	52.1	175.4
66.6	156.0	50.9	285.3
62.0	136.8	49.2	378.0

99.3-82.3mol% L_1 + L_2 + C (complex) 192.9°

Silver chloride (AgCl) + Indium chloride ($InCl_3$)

Vovkogon and Fialkov, 1945

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0.0	456	-	31.6	363	329
7.8	434	367	38.0	348	329
10.0	428	366	41.7	370	328
13.3	420	367	47.2	401	328
17.8	395	369	49.1	416	326
23.0	374	369	54.0	435	325
28.0	373	327	63.9	492	326
29.1	372	328	73.8	517	324

Beryllium chloride ($BeCl_2$) + Barium chloride ($BaCl_2$)

Schmidt, 1929

mol%	f.t.	mol%	f.t.
100	940	40	500
90	910	30	455
80	860	20	410
70	800	13	372
60	750	10	380
50	620	0	404
43	505		

Beryllium chloride ($BeCl_2$) + Lead chloride ($PbCl_2$)

Schmidt, 1929

mol%	f.t.	E	min
100	498	-	-
80	448	279	7.5
70	394	279	10
60	337	292	20
47.7	302	292	19
40	325	290	15
30	352	292	10
22.5	357	276	7.5
15	361	274	4
7.5	380	279	3
0	404	-	-

E: 53 mol% 292°

Beryllium chloride ($BeCl_2$) + Cadmium chloride ($CdCl_2$)

Schmidt, 1929

mol%	f.t.	E	tr.t.	min
100	570	-	-	-
80	549	316	-	5
69.3	524	322	252	5
60	498	326	273	7
50.5	466	326	286	5
38	426	326	293	8
30	398	326	298	7
20	346	327	300	9
10	355	340	278	9
5	372	318	273	7
0	404	-	300	-

E : 15 mol% 327°

Magnesium chloride ($MgCl_2$) + Calcium chloride ($CaCl_2$)

Menge, 1911

% by weight	% by analysis	f.t.	E	min.
100	100	777	-	-
95.2	97.9	773	601	1
90.4	94.9	773	606	1.9
80.4	85.3	754	599	5.5
71.1	72.2	715	606	8.5
61.2	64.1	695	609	17.0
51.3	55.0	666	622	17.0
42.2	45.2	-	621	39.7
31.1	32.1	645	620	21.0
20.8	22.1	666	614	13.1
10.5	10.9	696	607	6.4
5.3	6.9	698	591	4.6
0	0	711	-	-

Magnesium chloride ($MgCl_2$) + Strontium chloride ($SrCl_2$)

Sandonnini, 1912

mol%	f.t.	E	min.
0.0	712	-	-
6.5	700	532	-
13.6	675	535	20
28.7	610	536	50
34.2	585	535	90
42.1	-	535	110
49.4	E	535	140
59.8	680	536	100
73.8	770	535	60
85.2	825	535	30
100.0	860	-	-

Magnesium chloride ($MgCl_2$) + Barium chloride ($BaCl_2$)

Sandonnini, 1912

mol%	f.t.	E	min.	tr.t.
0.0	712	-	-	-
5.6	700	555	-	-
9.3	690	560	30	-
17.1	672	556	-	-
24.4	642	555	70	-
28.2	615	560	100	-
32.4	576	552	180	-
42.0	610	560	100	-
54.8	750	560	-	590
63.7	810	-	-	580
81.4	895	-	-	570
90.3	908	-	-	556
100.0	960	-	-	-

Magnesium chloride ($MgCl_2$) + Zinc chloride ($ZnCl_2$)

Menge, 1911

%	f.t.	E	min.
0	711	-	-
5	705	272	2.3
11.2	701	274	3.0
22.1	676	279	9.9
34.4	659	277	10.1
54.2	605	276	13.5
78.2	546	287	24.6
91.6	465	290	29.8
95	437	287	30.0
98	378	274	31.7
100	271	-	32.0

Magnesium chloride ($MgCl_2$) + Cadmium chloride ($CdCl_2$)

Menge, 1911

%	f.t.	m.t.	%	f.t.	m.t.
0.0	711	711	74.2	615	578
5.1	708	697	90.8	579	572
10.8	699	692	95.1	571	569
34.6	678	655	100	563	563
55.1	650	610			

Magnesium chloride ($MgCl_2$) + Stannous chloride ($SnCl_2$)

Menge, 1911

%	f.t.	E	min.
0	711	-	-
13.2	698	-	-
21.8	683	226	7.0
36.0	658	227	10.6
42.6	650	221	16.0
63.4	607	224	13.3
75.5	582	223	20.9
81.0	541	213	17.3
90	481	234	20.3
95	406	233	24.0
98	367	245	24.2
100	245	-	24.9

Magnesium chloride ($MgCl_2$) + Lead chloride ($PbCl_2$)

Menge, 1911

%	f.t.	E	min.
0	711	-	-
5.5	709	452	1.3
11.1	706	451	1.9
33.3	683	453	6.5
54.2	645	450	10.6
73.1	599	453	13.5
82.2	554	455	13.8
91.1	473	458	17.0
93.3	-	459	18.5
95.5	478	454	9.4
98.0	491	458	2.6
100	496	-	-

Magnesium chloride ($MgCl_2$) + Manganese chloride ($MnCl_2$)

Sandonnini, 1912

mol%	f.t.	mol%	f.t.
0.0	712	64.7	667
16.3	703	85.9	656
24.7	699	88.4	652
31.2	688	95.0	652
43.4	680	100.0	650
55.0	672		

Magnesium chloride ($MgCl_2$) + Ferrous chloride ($FeCl_2$)

Ferrari and Carugati, 1928

mol%	f.t.	mol%	f.t.
0	712	60	679
10	703	70	676
20	696	80	675
30	691	90	674
40	687	100	674
50	682		

Magnesium chloride ($MgCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari and Inganni, 1928

mol%	f.t.	mol%	f.t.
0	712	60	712
10	"	70	714
20	"	80	716
30	"	90	719
40	"	100	724
50	"		

Magnesium chloride ($MgCl_2$) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	190.2	77.9	207.6
90.5	188.6	72.5	224.2
87.6	187.4	70.9	227.4
84.6	186.4	69.1	350.0
82.1	188.0 (1+2)		

Calcium chloride ($CaCl_2$) + Strontium chloride ($SrCl_2$)

Sandonnini, 1911

mol%	f.t.	sat.t. of mixed crystals
0	772	-
10	724	-
20	700	535
25	671	550
30	662	560
35	646	560
40	668	550
50	695	535
60	730	-
70	771	-
90	842	-
100	872	-

Schaefer, 1914

%	f.t.	m.t.	sat.t. of mixed crystals	min.
100	870	-	-	-
89	827	812	-	-
76.92	776	755	-	-
72.62	762	732	493	40
67.71	743	721	503	80
58.82	701	679	525	120
58.77	670	654	541	140
37.97	661	658	541	90
32.25	673	665	528	60
26.31	705	671	491	40
13.69	739	708	-	-
0	773	-	-	-

Bukhalov and Bergman, 1952

mol%	f.t.	mol%	f.t.
0	773	44	679
5	764	48	697
10	753	55	716
15	739	60	736
20	715	70	770
25	697	80	808
30	675	90	840
35	656	100	868
40	665		

Min.: 35% 656°

Arndt and Geszler, 1908

t	d 0%	d 50%	d 100%
850	2.03	-	-
900	2.01	2.33	2.69
950	1.99	2.30	2.67
1000	1.97	2.28	2.645
1050	-	2.24	2.62

t	$\varkappa$ 0%	$\varkappa$ 50%	$\varkappa$ 100%
900	2.32	2.17	1.98
950	2.50	2.34	2.14
1000	2.66	2.47	2.29
1050	2.76	2.59	2.43

Calcium chloride ($CaCl_2$) + Barium chloride ($BaCl_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	780	50	695
10	750	60	760
20	700	70	815
30	650	80	870
36	590	90	910
40	625	100	960

Sandonnini, 1911

mol%	f.t.	E	min.	tr.t.
0	772	-	-	-
10	708	575	20	-
20	667	593	30	-
30	615	600	70	-
35	-	600	90	-
40	620	600	60	-
50	700	595	40	-
70	810	596	20	-
90	906	595	15	922
100	960	-	-	923

Schaefer, 1914

%	f.t.	E	min.	$tr.t._I$	min.	$tr.t._{II}$
0	773	-	-	-	-	-
13.20	743	591	20	-	-	-
24.87	712	595	160	-	-	-
35.26	679	596	210	-	-	-
44.57	644	602	330	-	-	-
50.26	621	602	380	-	-	-
55.57	-	602	440	-	-	-
60.55	619	602	230	-	-	-
65.23	646	592	30	631	150	-
69.64	687	592	-	629	145	-
73.79	722	-	-	629	120	-
84.91	822	-	-	622	50	-
97.27	936	-	-	611	20	919
100	960	-	-	-	-	924

Bukhalova and Bergman, 1951

mol%	f.t.	mol%	f.t.
0	773	43	616
10	730	49	626
20	686	52	626
25	663	55	683
28	645	58	729
31	626	61	751
34	614	67.5	800
37	597	85	838
40	607	100	962

(1+1) 629° E_1: 36.5 mol% 594°

E_2: 54 mol% 624°

Gromakov and Gromakova, 1953

mol%	f.t.	m.t.	
0	777	-	
10	718	578	
20	683	576	
30	632	579	
35	602	579	
38	580	E	(complex)
40	594	579	
45	621	584	
48.5	634	E	
50	650	584	
55	700	-	
60	741	-	

Calcium chloride ($CaCl_2$) + Manganese chloride ($MnCl_2$)

Sandonnini, 1911

mol%	f.t.	m.t.	sat.t. (mixed crystals)
0	772	-	-
10	738	-	-
20	704	-	-
30	660	-	-
40	640	620	-
50	610	597	-
55	597	590	-
60	589	583	462
65	583	583	470
70	590	585	460
80	616	605	-
90	636	624	-
100	650	-	-

Ferrari and Inganni, 1930

mol%	f.t.	decomposition of mixed crystals
0.0	782	-
10	739	-
19.3	710	-
30.0	674	-
42.3	635	-
50.0	619	-
60.0	596	475
65.4	590	474
70.0	592	475
80.6	608	473
90.0	627	-
100.0	650	-

Calcium chloride ($CaCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari and Inganni, 1930

mol%	f.t.	E	min.
0.0	782	-	-
9.5	744	-	-
19.3	709	611	270
25.4	690	612	300
40.6	629	616	615
45.7	614	614	930
51.0	621	614	795
59.6	642	-	-
65.0	651	614	486
69.1	658	-	-
75.0	674	613	195
86.8	684	-	-
85.0	693	613	105
88.2	700	-	-
100.0	722	-	-

Calcium chloride ($CaCl_2$) + Ferrous chloride ($FeCl_2$)

Ferrari and Inganni, 1930

mol%	f.t.	E	mol%	f.t.	E
0.0	782	-	66.6	617	-
10	758	562	69.2	624	-
19.5	734	588	75.0	637	593
30.0	696	590	80.0	650	-
39.9	658	594	85.01	660	593
50.0	617	594	90.0	667	-
55.5	592	592	100.0	674	-
60.0	600	592			

Calcium chloride ($CaCl_2$) + Zinc chloride ($ZnCl_2$)

Menge, 1911

%	f.t.	E	min.
0	777	-	-
8.2	754	255	4.4
18.4	726	241	7.0
29	690	243	11.2
49	621	271	13.1
70	534	291	14.8
90	403	271	22.7
95	349	291	26.1
98	294	291	23.3
100	274	-	28.3

Calcium chloride ($CaCl_2$) + Cadmium chloride ($CdCl_2$)

Menge, 1911

%	f.t.	m.t.	%	f.t.	m.t.
0	777	777	70	606	545
10	754	-	80	561	544
20	734	-	85	548	539
30	720	550	90	540	540
40	712	545	95	553	541
50	667	544	100	563	563
60	623	543			

Sandonnini, 1911

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	772	-	70	570	550
10	745	-	80	551	546
20	720	-	85	545	544
30	673	-	90	552	545
40	657	-	95	560	550
50	620	-	100	568	-
60	590	570			

Ferrari and Inganni, 1930

mol%	f.t.	decomposition of mixed crystals
0.0	782	-
10.0	735	-
20.2	704	-
30.0	681	-
39.1	652	-
50.0	617	414
55.0	605	415
59.92	592	412
65.0	578	414
70.0	566	413
79.9	548	414
85.0	537	413
90.0	545	414
95.51	554	412
100.0	564	-

Calcium chloride ($CaCl_2$) + Stannous chloride ($SnCl_2$)

Menge, 1911

%	f.t.	E	min.
0	777	-	-
66.7	644	238	16.9
72.8	634	239	18.3
78.7	580	242	21.4
89.6	446	244	24.9
94.7	344	241	25.0
97.9	310	241	25.2
100	245	-	25.6

Calcium chloride ($CaCl_2$) + Lead chloride ($PbCl_2$)

Menge, 1911

%	f.t.	E	min.
0	777	-	-
5	766-773	-	-
10	756-763	-	-
20	748	448	3.9
30	717-727	447	3.5
50	704	457	5.7
70	587	463	9.8
80	-	463	10.2
83	-	473	12.5
87	480	468	8.7
90	472	467	4.3
95	483	465	2.6
100	496	-	-

Sandonnini, 1911

mol%	f.t.	E	min.
0	772	-	-
10	730	458	-
20	700	465	15
30	670	465	25
40	630	467	40
50	608	467	80
60	570	466	100
65	545	465	130
70	500	468	150
80	490	467	170
85	477	467	60
90	485	468	15
95	489	470	-
100	495	-	-

Strontium chloride ($SrCl_2$) + Barium chloride ($BaCl_2$)

Sandonnini, 1911

mol%	f.t.	tr.t.	mol%	f.t.	tr.t.
0	872	-	40	856	-
10	864	-	60	893	795
25	862	-	70	914	821
30	859	-	90	945	885
35	854	-	100	960	923

Vortisch, 1914

%	f.t.	tr.t.	%	f.t.	tr.t.
0	870	-	66.34	870	750
12.74	859	(635)	79.76	889	816
30.45	848	(650)	92.20	923	877
46.69	852	682	100	955	922
56.78	858	716			

Strontium chloride ($SrCl_2$) + Zinc chloride ($ZnCl_2$)

Sandonnini, 1912

mol %	f.t.	E	min.
100	275	-	-
96.5	310	270	-
90.0	350	270	-
81.3	422	250	-
73.4	450	235	-
62.1	470	220	-
55.4	473	-	-
48.4	-	476	-
46.2	-	477	190
41.6	580	476	150
37.8	620	477	100
35.9	650	476	80
18.9	752	477	60
0	860 (1+1)	476	40

Strontium chloride ($SrCl_2$) + Cadmium chloride ($CdCl_2$)

Sandonnini, 1911

mol%	f.t.	E	min.
0	872	-	-
10	803	493	20
20	730	498	40
30	660	500	70
40	590	502	100
50	530	503	120
60	-	503	150
70	525	500	70
80	532	500	40
90	546	489	20
100	568	-	-

Srontium chloride ($SrCl_2$) + Lead chloride ($PbCl_2$)

Sandonnini, 1911

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	872	-	70	560	523
10	830	-	80	540	510
20	792	-	90	525	505
40	715	-	95	510	500
50	670	-	100	495	-
60	610	-			

Strontium chloride ($SrCl_2$) + Manganese chloride ($MnCl_2$)

Sandonnini, 1911

mol%	f.t.	E	min.
0	872	-	-
10	800	489	20
20	750	498	50
30	630	499	80
40	547	499	100
45	-	499	120
50	509	499	100
60	523	499	70
70	573	496	50
80	593	497	40
90	626	497	20
95	637	496	-
100	650	-	-

Strontium chloride ($SrCl_2$) + Ferrous chloride ($FeCl_2$)

Ferrari and Inganni, 1930

mol%	f.t.	E	mol%	f.t.	E
0	877	-	60	562	539
6.5	814	542	70	587	540
20	740	538	80	622	540
30	669	542	90	646	540
40	608	542	100	680	540
50	541	541			

Strontium chloride ($SrCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari and Inganni, 1930

mol%	f.t.	E	mol%	f.t.	E
0	877	-	49.9	587	557
9.4	808	557	58.8	613	556
19.4	730	559	71.5	644	556
30.1	658	561	81.1	674	556
40.5	564	564	100	722	-

Barium chloride ($BaCl_2$) + Zinc chloride ($ZnCl_2$)

Sandonnini, 1912

mol%	f.t.	E	tr.t.	min.
100	275	-	-	-
96.3	356	272	-	-
88.6	392	240	-	-
82.0	415	242	-	-
74.5	462	240	-	-
61.0	468	-	-	-
57.9	470	-	-	-
49.2	535	-	470	120 (1+1)
46.6	645	-	"	80
42.0	710	-	"	80
34.3	830	-	471	70
19.3	830	-	470	60
8.2	904	-	470	40
7.0	918	-	923	30
3.8	936	-	923	-
0	960	-	-	-

Barium chloride ($BaCl_2$) + Cadmium chloride ($CdCl_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	590	50	560
90	565	40	650
80	540	30	740
70	500	20	815
65	480	10	860
60	500	0	960

Sandonnini, 1911

mol%	f.t.	E	min.
0	960	-	-
3	935	427	-
5	910	430	-
10	880	447	20
20	825	451	40
30	750	445	40
40	650	446	50
50	470	450	70
60	454	450	150
70	483	449	60
80	520	448	30
90	547	448	20
95	555	443	-
100	568	-	-

Barium chloride ($BaCl_2$) + Stannous chloride ($SnCl_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
	960°		
100	890	40	895
90	880	30	910
80	875	20	925
70	870	10	940
60	875	0	960
50	885		

Barium chloride ($BaCl_2$) + Lead chloride ($PbCl_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	515	75	570
90	535	70	590
80	560	60	650
		0	900

Sandonnini, 1911

mol%	f.t.	m.t.	tr.t.
0	960	-	923
3	940	-	912
5	928	-	913
10	906	-	-
20	860	-	-
30	800	-	-
40	750	-	-
50	700	-	-
60	620	-	-
70	570	525	-
80	548	512	-
90	514	500	-
100	495	-	-

Lorenz, Frei and Jabs, 1908

$d= 3.46-8/30000^t + (2.04-1/1000^t) \,.\, A/100$

where A= mol% Lead chloride

Barium chloride ($BaCl_2$) + Manganese chloride ($MnCl_2$)

Sandonnini, 1912

mol%	f.t.	E	tr.t.	min.
0	960	-	923	-
3	943	540	923	-
10	904	-	500	-
20	853	533	-	-
30	770	534	497	30
33	720	554	493	-
40	667	554	496	-
45	-	534	495	40
50	554	-	498	50
55	532	-	505	100
60	520	-	506	100
65	-	-	503	120 (1+2)
70	556	-	508	100
80	595	-	504	80
90	623	-	498	50
95	693	-	495	10
100	650	-	-	-

Barium chloride ($BaCl_2$) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol %	f.t.	mol %	f.t.
100	190.2	77.2	209.4
86.0	192.5	72.9	255.5
79.9	198.8		

98.0-87.0 mol % L_1+L_2+C (1+2) 191.5°

Zinc chloride ($ZnCl_2$) + Cadmium chloride ($CdCl_2$)

Herrmann, 1911

%	f.t.	E	min.
100	568	-	-
90	553	262	2.4
50	470	262	7
30	387	262	8
10	288	262	11
5	269.5	262	12
0	261.5	-	-

Zinc chloride ($ZnCl_2$) + Stannous chloride ($SnCl_2$)

Herrmann, 1911

%	f.t.	E	min.
0	261.5	-	-
10	260	172	2.5
30	248	171	9
50	217	171	15
70	190	180	16
90	220	170	5
100	247.2	-	-

Zinc chloride ($ZnCl_2$) + Lead chloride ($PbCl_2$)

Muromtsev and Nazarova, 1946

mol%				
L	V	p	p_1	p_2
		700°		
100	-	26.5	-	26.5
85.24	26.66	83	61	22.2
48.04	6.44	235	220	15.1
25.10	2.75	350	340	9.6
0	-	502	502	-

Herrmann, 1911

%	f.t.	E	min.
100	501	-	-
90	472	243	1.6
80	414	260	6.5
70	390	260	6.3
60	355	262	8
50	332	256	8.1
40	321	265	9.3
30	304	268	13
20	285	268	11
10	273	-	17
1	262	-	14
0	261.5	-	

Urazov and Sokolova, 1944

mol %	f.t.	E	tr.t.
0	320	-	308
0.3	320	-	-
0.5	318	-	-
1.0	318	-	-
2.0	318	-	-
3.0	320	-	-
3.5	322	-	-
5.0	324	-	304
6.5	324	-	-
8.0	322	-	-
9.0	318	-	303
20	320	-	308
30	322	-	317
40	330	318	323
50	344	318	330
60	372	319	362
70	410	-	411
80	445	-	445
90	480	-	500
100	500	-	-

Ogai and Chatilo, 1949 (fig.)

mol%	f.t.	mol%	f.t.
0	315	60	365
20	308	80	460
40	325	100	500

Wachter and Hildebrand, 1930

t	d	t	d
	100 %		
545.3	4.872	526.7	4.902
537.2	4.886	514.4	4.924
535.5	4.886	502.1	4.944
	49.8 %		
552.5	3.703	535.2	3.717
545.4	3.709	510.2	3.733
	0 %		
550.7	2.401	503.2	2.422
549.3	2.401	493.7	2.427
524.5	2.413	485.7	2.431
509.3	2.420		

t	e	t	e
	100 %		
499.1	1.2731	544.8	1.2445
510.5	.2653	549.5	.2417
517.5	.2631	554.8	.2394
521.7	.2603	556.5	.2374
525.3	.2563	560.3	.2350
529.2	.2554	566.0	.2310
536.8	.2507	572.0	.2281
539.4	.2488	578.5	.2249
541.8	.2469	582.1	.2227
	85.5 %		
516.6	1.2703	536.7	1.2574
520.7	.2670	540.6	.2549
527.6	.2629	554.6	.2461
531.3	.2609	587.0	.2248
	68.8 %		
495.8	1.2934	551.0	1.2546
512.7	.2828	560.0	.2546
525.2	.2753	566.2	.2510
542.5	.2650	570.0	.2492
	59.5 %		
493.4	1.3021	547.5	1.2713
512.6	.2921	557.9	.2650
534.4	.2792	561.3	.2636
538.0	.2770	563.0	.2617
545.3	.2728	568.2	.2584
	49.0 %		
495.0	1.3125	562.3	1.2738
504.8	.3068	567.4	.2708
528.2	.2933	595.5	.2540
537.9	.2882	602.2	.2503
	30.1 %		
494.7	1.3341	529.0	1.3162
503.4	.3296	541.3	.3098
512.3	.3253	545.7	.3079
517.0	.3224	579.1	.2910
525.0	.3185		

Zinc chloride ($ZnCl_2$)+ Manganese chloride ($MnCl_2$)

Sandonnini, 1912

mol%	f.t.	E	min.
0.0	275	-	-
5.4	295	275	120
12.0	350	276	90
22.7	390	276	80
30.5	452	267	70
52.3	496	265	60
57.3	530	250	60
70.5	566	230	-
80.5	590	240	-
90.2	610	230	-
100.0	650	-	-

Zinc chloride ($ZnCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari and Inganni, 1930

mol%	f.t.	E	min.
0	300	-	-
10	-	300	495
20	452	"	"
30	500	"	450
40	558	299	405
50	598	292	360
60	630	289	420
70	650	290	330
80	680	281	180
90	700	290	150
100	725	-	-

Bassett and Bedwell, 1931

%	f.t.	%	E
7.17	300	7.72	311.0
10.37	367	16.42	311.4
18.14	480	19.45	312.8
26.95	550	27.03	312.0
30.77	580	37.48	310.7
38.48	605	41.44	310.8
53.99	627		
67.72	650		
73.41	682		
95.93	720		
96.68	725		
98.18	730		
100	735		

Zinc chloride ($ZnCl_2$) + Ferric chloride ($FeCl_3$)

Herrmann, 1911

%	f.t.	m.t.	E	min.
100	298	-	-	-
90	-	274	-	-
87	-	242	-	-
82	275	-	205	2
70	263	-	207	7.5
50	242	-	211	16
30	-	-	214	26
20	224	-	213	14
10	244	-	216	4
5	-	225	-	-
0	261.5	-	-	-

Zinc chloride ($ZnCl_2$) + Indium chloride ($InCl_3$)

Vovkogon and Fialkov, 1945

mol%	f.t.	E	mol%	f.t.	E
0.0	340	-	17.6	416	266
0.8	320	-	18.0	410	272
2.2	280	-	20.8	426	272
3.2	298	273	23.2	440	268
4.0	300	276	23.8	438	271
4.6	302	276	27.2	448	260
6.0	328	272	28.4	464	261
7.2	350	272	36.8	490	263
9.2	364	272	39.2	480	270
11.0	368	276	40.8	496	268
12.4	393	268	50.4	516	268
12.8	380	270	51.6	516	270
13.2	400	262	67.1	528	-
14.2	390	268	75.2	550	-

Zinc chloride ($ZnCl_2$) + Bismuth chloride ($BiCl_3$)

Herrmann, 1911

%	f.t.	min.	E	min.
100	224	-	-	-
97.5	217	-	207	9
95	-	-	215	17
90	246	-	215	16
85	273	1.5	204	15
70	274	4.6	215	13
50	280	7.1	204	10
30	285.5	12	206	6
20	280	13	203	4
10	280	14 $L_1 + L_2$	-	-
5	281	17	204	-
2.5	280	267	-	-
0	261.5	-	-	-

Jänecke, 1933

%	f.t.	E
60	264	215-210
69	263	207.2
55	267	198
57	276	207.2

Cadmium chloride ($CdCl_2$) + Stannous chloride ($SnCl_2$)

Herrmann, 1911

%	f.t.	E	min.
0	568	-	-
10	552	229	1.5
50	463	230	9.5
70	380	230	11.6
90	-	229	14.0
95	236	230	6.0
100	247.2	-	-

Sandonnini and Scarpa, 1912

mol%	f.t.	E	min.
0	560	-	-
5	560	233	-
10	550	233	40
20	525	231	50
30	507	233	80
40	478	235	100
50	441	238	120
60	415	232	160
70	360	235	170
80	286	233	180
85	270	235	200
90	E	233	210
95	240	233	100
100	250	-	-

Cadmium chloride ($CdCl_2$) + Lead chloride ($PbCl_2$)

Herrmann, 1911

%	f.t.	E	min.
100	501	-	-
90	461	385	3.2
80	408	385	7.4
70	-	385	11.5
60	422	385	8.4
50	452.5	379	7.0
40	490	380	6.3
30	512	378	4.5
20	529	378	3.0
10	554	374	1.7
0	568	-	-

Sandonnini, 1912

mol%	f.t.	E	min.
0	568	-	-
5	555	384	-
10	543	385	20
20	524	385	30
30	510	388	50
40	472	390	60
50	428	390	80
60	400	389	90
65	E	389	120
70	405	389	60
80	440	389	30
90	460	385	20
95	485	385	-
100	495	-	-

Ilyasov, Bostandzhiyan and Bergman, 1956.

mol%	f.t.	mol%	f.t.
0.0	568	42.5	460 tr.t.
2.5	562	45.0	453
5.0	557	50.0	437
7.5	553	55.0	419
10.0	549	60.0	399
15.0	539	63.5	385 E
20.0	525	65.0	392
25.0	517	67.5	399
30.0	505	70.0	408 tr.t.
35.0	488	72.5	419
37.5	480	75.0	429

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ		mol%	σ	
	600°	700°		600°	700°
0	83.5	81	60	107	101.5
20	90	86.5	80	116	108
40	99	94.5	100	125	116

Boardman,Dorman and Heyman, 1949

mol%	molar vol. in cc
	600°
100	57.90
79.4	56.90
67.2	56.30
41.8	55.50
20.1	54.81
0	54.46

Cadmium chloride ($CdCl_2$) + Ferrous chloride ($FeCl_2$)

Ferrari and Carugati, 1928

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	560	-	60	607	581
10	560	-	70	621	598
20	562	-	80	641	621
30	571	-	90	658	648
40	580	-	100	674	-
50	591	568			

Cadmium chloride ($CdCl_2$) + Manganese chloride ($MnCl_2$)

Sandonnini and Scarpa, 1911

mol %	f.t.	m.t.	mol %	f.t.	m.t.
100	650	-	40	591	577
90	640	630	30	583	574
80	624	609	20	573	570
70	617	603	10	570	568
60	606	589	0	568	-
50	601	587			

Cadmium chloride ($CdCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari and Inganni, 1930

mol%	f.t.	mol%	f.t.	m.t.
0	560	66.5	643	596
5.0	561	74.5	663	622
12.0	561	85.5	688	665
15.0	562	87.5	694	674
27.5	568	94.5	710	700
45.0	593	100.0	724	-
54.0	614			

Cadmium chloride ($CdCl_2$) + Indium chloride ($InCl_3$)

Vovkogon and Fialkov, 1945

mol%	f.t.	E	mol%	f.t.	E
0.0	569	-	49.6	479	477
5.2	564	-	51.2	487	475
13.4	554	473	53.8	498	475
14.5	552	-	59.0	514	477
19.3	546	473	64.4	527	476
24.8	534	474	69.6	538	476
30.1	526	474	76.8	551	473
38.2	507	476	88.7	563	-
44.6	493	475	100.0	585	-
48.8	479	476			

E : 49% 476°

Mercuric chloride ($HgCl_2$) + Lead chloride ($PbCl_2$)

van Driel, 1935

mol%	f.t.	L_1+L_2+C	sat.t.	
0.00	281	-	-	-
1.8	363	-	-	-
4.3	402	-	-	-
10.2	-	415	-	-
10.5	-	-	510	-
14.9	-	414	-	-
15.2	-	-	509	-
21.0	-	414	-	-
21.5	-	-	508	-
29.3	-	413	-	-
30.9	-	-	544	-
40.2	-	413	-	-
41.0	-	-	-	474
43.5	-	415	-	-
43.5	-	-	-	419
48.3	416	-	-	-
60.2	422	-	-	-
71.2	436	-	-	-
80.2	450	-	-	-
90.5	474	-	-	-
100.00	499	-	-	-

E:281°

Mercuric chloride ($HgCl_2$) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	190.2	51.8	151.3
97.0	190.0	44.1	192.0
92.3	189.7	41.9	200.5
81.5	189.0	30.8	216.2
73.6	184.6	29.1	238.1
67.2	177.5	17.4	259.7
62.6	167.5	10	272.1
58.6	167.4		

Mercuric chloride ($HgCl_2$) + Antimonium chloride ($SbCl_3$)

Kendall, Crittenden and Miller, 1923

%	f.t.	%	f.t.
100	73.4	74.6	183.5
98.0	72.6	55.7	218.8
94.8	85.4	50.7	225.6
93.5	101.7	41.6	234.5
91.1	114.8	30.4	246.4
90.2	126.3	20.4	257.1
86.9	142.1	13.9	264.2

Stannous chloride ($SnCl_2$)+ Lead chloride ($PbCl_2$)

Sandonnini and Scarpa, 1911

mol%-	f.t.	m.t.	mol%	f.t.	m.t.
0	250	-	50.3	403	393
7.06	270	259	61.3	430	420
14.6	298	287	73.1	450	440
22.7	328	319	85.8	473	461
31.0	350	339	100	495	-
40.7	378	378			

Herrmann, 1911

%	f.t.	m.t..	%	f.t.	m.t.
100	501	501	40	356	285
90	471	436	30	334.7	255
80	452	395	20	302.5	243
70	429	357	10	275	241
60	403.5	324	0	247.2	247.2
50	379	309			

Benrath and Tesche, 1920

t	$\varkappa \cdot 10^6$					
	0%	10%	20%	30%	40%	50%
100	0.3	1.1	1.0	0.9	0.7	0.5
10	0.4	1.3	1.2	1.3	1.1	0.8
20	0.6	1.7	1.5	1.4	1.4	1.0
30	1.0	2.2	2.1	1.9	1.7	1.4
40	1.3	2.9	2.6	2.5	2.2	1.8
50	1.9	3.7	3.4	3.2	2.8	2.3
60	3.0	4.8	4.6	4.2	3.6	3.0
70	5.9	6.2	5.6	5.0	4.2 °	3.6
80	9.2	8.3	7.2	6.2	5.6	4.2
90	14.0	11.0	9.6	7.7	6.6	5.4
200	25.0	19.5	15.0	9.9	8.9	6.5
10	46.1	28.7	19	13.4	9.8	7.8
20	70.4	42.0	32	20.1	12.2	9.0
30	130	79.6	54	-	15.1	10.6
40	200	140	80	40.4	1.9.5	13.0
50	3117	142	94	51.3	24.6	16.6
60	-	-	142	80.0	32.5	21.5
70	-	-	-	205	50.7	28.0
80	-	-	-	775	64.5	37.7
90	-	-	-	-	255	54.0
300	-	-	-	-	-	95.5
10	-	-	-	-	-	143.3
20	-	-	-	-	-	317

t	$\varkappa \cdot 10^{6}$ 60%	70%	80%	90%	100%
100	0.2	-	-	-	-
10	0.4	0.2	0.05	-	-
20	0.6	0.3	0.1	-	-
30	0.9	0.5	0.2	-	-
40	1.3	0.7	0.4	-	-
50	1.7	1;1	0.5	0.1	-
60	2.3	1.5	0.7	0.3	0.03
70	2.7	1.8	0.9	0.4	0.09
80	3.2	2.3	1.2	0.6	0.16
90	4.0	2.8	1.6	0.9	0.2
200	5.0	3.5	2.0	1.2	0.3
10	6.0	4.3	2.6	1.4	0.5
20	7.5	5.2	3.2	1.8	0.6
30	9.0	6.4	4.3	2.2	0.9
40	10.9	7.2	5.1	2.8	1.2
50	13.9	9.9	6.8	3.6	1.4
60	16.8	12.6	8.4	4.8	2.0
70	21.0	15.4	10.3	5.8	2.3
80	26.1	18.8	13.8	7.5	3.8
90	33.4	22.3	15.8	9.0	5.0
300	43.5	25.4	18.1	10.7	5.6
10	53.3	32.0	22.1	12.7	9.6
20	82.6	40.4	28.0	15.6	12.0
30	131	55.1	36.8	19.0	15.3
40	-	67.4	46.8	25.0	19.4
50	-	90.0	67.9	38.7	23.0
60	-	136.3	91.9	50.3	23.0
60	-	-	149	79.5	30.8
70	-	-	176.4	100.0	39.2
80	-	-	233	141.1	51.6
90	-	-	-	175.3	63.1
400	-	-	-	231	77.6
10	-	-	-	-	93.3
20	-	-	-	-	109.7
30	-	-	-	-	134.6
40	-	-	-	-	183.9
50	-	-	-	-	222.7
60	-	-	-	-	442.1
70	-	-	-	-	305.9

Stannous chloride ($SnCl_2$) + Manganese chloride ($MnCl_2$)

Sandonnini and Scarpa, 1911

mol%	f.t.	E	min.
100	850	-	-
90	630	225	30
80	607	224	50
70	582	224	70
60	549	225	80
50	516	230	90
40	478	230	100
30	423	231	120
20	370	231	130
10	300	232	150
5	E	233	200
2	240	231	30
0	250	-	-

Stannous chloride ($SnCl_2$) + Ferrous chloride ($FeCl_2$)

Kuromtsev and Nazarova, 1946

mol% V	L	p	p_1	p_2
		500°		
0	0	314	314	-
1.19	15.79	270	267	3.3
2.13	34.96	200	196	4.5
2.74	43.18	169	164	4.8
6.87	54.81	121	113	8.3

Ferrari and Colla, 1933

mol%	f.t.	E	mol%	f.t.	E
100	675	-	27.75	487	240
88.52	652	237	20.90	449	240
82.84	640	237	12.26	380	242
74.90	624	232	8.24	322	242
63.14	592	234	4.34	362	244
51.86	568	238	3.38	-	242
43.00	538	240	0	248	-

Stannous chloride ($SnCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari and Colla, 1933

mol%	f.t.	E	mol%	f.t.	E
100	724	-	44.52	539	240
94.21	700	240	30.40	486	243
87.60	682	240	14.13	377	243
81.76	668	236	6.20	-	241
74.08	641	240	2.50	-	240
64.40	612	240	0	248	-
57.39	584	240			

Stannous chloride ($SnCl_2$) + Antimonium chloride ($SbCl_3$)

Kendall, Crittenden and Miller, 1923

%	f.t.	%	f.t.
100	73.4	7.3	240.7
99.90	74.9compl.	5.3	241.3
99.29	169.0	2.1	244.5
99.15	174.9	0	246.8
8.3	242.1		

98.6 - 8.8mol% $C + L_1 + L_2$

Stannous chloride ($SnCl_2$) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	246.8	51.5	158.7
99.02	191.0 compl.	46.1	136.9
98.89	191.3	49.3	158.2 (1+1)
84.7	191.0	45.4	153.3
81.9	188.2	36.5	142.3
77.9	187.0 (1+2)	30.1	135.4
71.5	204.4	31.5	138.3
69.3	207.4	29.1	154.3
66.5	209.1	25.0	176.3
61.2	195.0	13.9	223.4
56.9	178.6	9.2	235.7
52.9	164.1	0	246.8

98.5 -85.7mol% $L_1 + L_2$ + C (complex) 192.0°

Lead chloride ($PbCl_2$) + Manganese chloride ($MnCl_2$)

Sandonnini and Scarpa, 1911

mol%	f.t.	E	min.
100	650	-	-
95	635	395	-
90	627	401	30
80	600	403	50
70	574	400	60
60	540	408	70
50	510	408	90
40	463	408	110
30	-	408	160
20	429	408	70
10	460	406	50
5	485	406	20
0	495	-	-

Lead chloride ($PbCl_2$) + Ferrous chloride ($FeCl_2$)

Herrmann, 1911

%	f.t.	E	min.
0	501	-	-
10	482	176.5	3.8
20	441	178	5.6
30	377	178	9.3
40	268.5	176	10.8
50	-	177	15.0
60	212	181	10.0
70	256	179	8.2
80	283	181	7.0
90	293	181	4.6
100	298	-	-

Ferrari and Colla, 1933

mol%	f.t.	E	mol%	f.t.	E
100	675	-	37.54	489	417
90.22	651	406	33.77	461	421
82.30	627	400	23.48	441	422
72.40	592	405	19.13	450	421
64.44	566	413	8.53	473	418
52.04	531	416	0	489	-

Lead chloride ($PbCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari and Colla, 1933

mol%	f.t.	E	mol%	f.t.	E
100	724	-	28.62	475	425
93.95	668	421	23.74	444	424
72.27	642	424	20.89	439	424
61.98	601	421	17.20	452	424
54.85	573	424	11.41	464	424
40.55	521	424	0	489	-

Lead chloride($PbCl_2$) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

%	f.t.	%	f.t.
100	190.2	67.5	272.5
99.37	191.9	63.0	266.9
83.1	211.9	61.0	257.0
79.9	220.4	56.4	241.6
76.3	235.5	54.5	234.9
72.0	253.5	47.7	268.4
69.5	267.7	42.3	296.8

Lead chloride ($PbCl$) + Indium chloride ($InCl_3$)

Vovkogon and Fialkov, 1945

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	500	-	31.2	388	374
4.4	488	368	37.0	420	374
7.6	479	369	48.8	455	372
11.6	463	373	51.8	467	370
14.4	459	372	61.8	494	369
15.7	454	374	73.0	523	-
22.6	416	371	82.1	551	-
26.3	375	-	89.8	567	-
27.4	373	-			

E : 28% 374°

Lead chloride ($PbCl_2$) + Bismuth chloride ($BiCl_3$)

Herrmann, 1911

%	f.t.	tr.t.	min.	E	min.
0	501	-	-	-	-
10	479	324	2	212	-
15	465	323	3	210	0.5
20	455	325	2.8	206	2
30	427	327	2.6	207.5	4.5
40	400	323	1.8	210	6
50	367	322	1.3	214	10
60	344	321	1.7	214	10.8
70	325	-	-	212	12
80	288	-	-	211	15
90	-	-	-	219	16.6
95	220	-	-	212	11
100	224	-	-	-	-

Manganese chloride ($MnCl_2$) + Ferrous chloride ($FeCl_2$)

Ferrari, Celeri and Giorgi, 1929

mol%	f.t.	mol%-	f.t.
0	650	60	661
10	651	70	665
20	652	80	668
30	654	90	671
40	656	100	674
50	659		

Manganese chloride ($MnCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari and Inganni 1930

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	649	-	49.5	665	653
5.0	650	-	60.0	674	660
10.0	650	-	69.5	684	670
20.0	651	-	80.5	696	683
30.0	653	-	89.5	706	700
40.0	659	-	100.0	724	-

Manganese chloride ($MnCl_2$) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	190.2	79.1	203.9
91.6	190.0	76.5	212.9
90.3	189.0	68.6	226.9
86.5	186.4	67.6	271.1
83.9	185.4	65.1	308.8
80.9	197.2 (1+2)		

Ferrous chloride ($FeCl_2$) + Cobalt chloride ($CoCl_2$)

Ferrari, Celeri and Giorgi, 1929

mol%	f.t.	mol%	f.t.
100	722	40	689
90	716	30	685
80	712	20	681
70	707	10	677
60	700	0	674
50	694		

Ferrous chloride ($FeCl_2$) + Ferric chloride ($FeCl_3$)

Schäffer and Bayer, 1953

mol%	f.t.	mol%	f.t.
100	307.5	83	340
95	305	81	365
90	301.5	79	387
86.5	297.5 E	77	405
85	315		

Aluminum chloride ($AlCl_3$) + Ferric chloride ($FeCl_3$)

Morozov, 1956 (fig.)

%	f.t.	%	f.t.
0	194	50	280
20	245	60	285
30	260	80	290
40	270	100	303

C %	V	t	p_1	p_2	p
25	2.89	150	54.07	1.32	55.4
50	traces	150	39.3	1.6	41.0
75	4.66	150	26.5	1.5	28.0
25	2.81	183	370.7	8.79	379.5
50	2.54	183	266.7	5.7	272.4
75	5.77	183	162.9	8.2	171.1

Aluminum chloride ($AlCl_3$) + Antimonium chloride ($SbCl_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
0	190.2	59.5	147.6
9.6	190.0	64.2	137.3
17.1	188.6	74.4	112.9
21.9	186.9	77.0	106.1
32.2	181.6	84.8	83.0
41.4	173.6	92.5	89.0
46.7	168.4	95.9	71.1
52.2	160.7	100.0	73.4

Aluminum chloride ($AlCl_3$) + Stannic chloride ($SnCl_4$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
0	190.2	83.4	139.5
6.8	188.6	89.6	130.9
19.3	182.6	93.2	122.0
34.8	174.2	95.9	113.4
47.3	167.8	98.8	89.9
53.8	164.5	99.6	65.7
63.2	158.3	100	30.2
74.4	150.0		

Aluminum chloride($AlCl_3$) + Tantalum pentachloride ($TaCl_5$)

Morozov, 1956 (fig.)

C %	V	p_1	p_2	p
		194°		
68.2	8.0	694.4	43.6	738.1
56.1	22.7	336.6	73.4	410.0
76.2	63.9	96.0	126.5	222.5

Morozov, Korshunov and Simonich, 1956.

%	f.t.	E	%	f.t.	E
100.0	220	-	38.8	151	116
90.6	197	106.6	27.4	160	113
81.3	162	107	21.8	167	98
70.7	131	117	13.2	176	-
63.0	-	117	0	194	-

E : 63% 117°

L %	V	p_1	p_2	p
		155°		
100.0	100.0	-	21	21
75.7	36.4	49.5	20	70
50.8	23.2	86.3	19.5	105.8
25.0	18.2	104.9	17.4	122.3
0	0	125.0	-	125.0
		175°		
100.0	100.0	-	80.0	80.0
75.4	62.8	75.0	73.5	148.0
54.5	10.0	316.0	35.3	341.0
25.7	8.4	416.0	28.3	444.3
0	0	450.0	-	450.0

Aluminum chloride ($AlCl_3$) + Niobium pentachloride ($NbCl_5$)

Morozov, 1956 fig.

C %	V	t	p_1	p_2	p
100	100	182	-	85	85
70	69	182	30	57	37
65	40	182	80	45	125
48	15.7	182	185	30	215
25	5.8	182	500	15	515
100	100	205	-	250	250
70	73.2	205	50	135	185
65	45.5	205	135	110	245
48	29.0	205	285	85	370

Morozov, Korshunov and Simovich, 1956.

%	f.t.	E	%	f.t.	E
0	194	-	68.7	134	114
15.0	188	114	69.0	134	114
25.5	176	119	74.5	148	120
38.0	168	126	79.7	176	114
51.0	146	126	91.3	196	-
59.0	133	128	100	203	-
65.0	125	-			

E: 65% 125°

% L	% V	p_1	p_2	p
		150°		
100.0	100.0	-	20	20
70.0	53.0	22	19	41
65.0	23.0	40	18	58
48.0	14.0	62	12	74
25.0	6.2	75	10	85
0.0	0.0	100.0	-	100.0
		170°		
100.0	100.0	-	50	50
70.0	68.0	19	40	59
65.0	38.0	55	25	80
48.0	8.5	145	15	100
25.0	6.6	275	10	285
0.0	0.0	400	-	400

Antimonium chloride ($SbCl_3$) + Bismuth chloride ($BiCl_3$)

Starokadomokaya, 1939

mol %	f.t.	mol %	f.t.
5.36	86	30.79.	165
6.04	96	53.61	188
15.84	135	63.78	201
28.19	163		

Antimonium chloride ($SbCl_3$) + Stannic chloride ($SnCl_4$)

Kendall, Crittenden and Miller, 1923

%	f.t.	%	f.t.
0	73.4	45.2	65.9
4.6	70.9	52.9	66.0
6.2	69.4	52.9	65.4
13.5	67.7	61.1	65.9
14.6	67.4	61.1	64.5
17.8	67.0	64.8	65.8
19.1	66.8	64.8	69.0
23.9	66.5	99.7	-30.5
35.8	66.3	99.7	-30.5
35.8	62.9	100	-30.2
45.2	66.1		

35.8 - 73.5mol% L_1 + L_2

Antimony trichloride ($SbCl_3$) + Antimony pentachloride ($SbCl_5$)

Aten, 1909

mol %	f.t.	mol %	f.t.
100	4.0	44.7	59.0
96	2.5	28.3	62.5
89.6	24.0	28.9	65.0
82.6	39.0	25.8	66.0
72.5	48.5	7.5	70.0
60.2	54.0	0	73.0
49.5	57.0		

Bismuth chloride ($BiCl_3$) + Ferric chloride ($FeCl_3$)

Herrmann, 1911

%	f.t.	E	min.
0	224	-	-
5	216	168.5	4.7
10	201	173.5	8.6
23	-	171.5	25
30	209	172	23
50	252	171	16
75	285	170	8
95	293	170	2.2
100	298	-	-

Ferric chloride ($FeCl_3$) + Tantalum pentachloride ($TaCl_5$)

Morozov, 1956 (fig.)

%	f.t.	E	%	f.t.	E
0	303°	-	50	280	200
20	295	200	70	270	200
30	285	195	87.3	200	200
40	283	200	100	220	-

t	% C	% V	p_1	p_2	p
150	49.0	100.0	0.0	16.6	16.6
175	49.0	100.0	0.0	54.6	54.6
199	49.0	98.0	4.5	194.2	198.7
214	49.0	97.5	8.5	291.0	299.5
127	49.0	100.0	5.6	0.0	5.6
137	49.0	100.0	7.4	0.0	7.4
153	49.0	100.0	23.8	0.0	23.8
157	49.0	100.0	31.8	0.0	31.8
175	49.0	100.0	74.8	0.0	74.8
182	49.0	99.5	161.0	0.8	161.8
185	49.0	99.5	168.0	0.8	168.8

Ferric chloride ($FeCl_3$) + Niobium pentachloride ($NbCl_5$)

Morozov, 1956 (fig.)

%	f.t.	E	%	f.t.	E
0	303	-	60	270	198
20	300	198	80	220	198
30	295	198	86.5	198	198
40	285	198	100	203	-
50	280	198			

% C	% V	t	p_1	p_2
50.0	100.0	147	0.0	10.0
50.0	100.0	153	0.0	27.5
50.0	100.0	171	0.0	26.5
50.0	100.0	180	0.0	82.3
50.0	99.5	196	0.8	168.0
50.0	98.5	200	2.5	179.0

Titanium chloride ($TiCl_4$) + Stannic chloride ($SnCl_4$)

Nasu, 1933

mol %	f.t.	mol %	f.t.
0.0	-24.8	58.0	-32.2
5.3	-26.0	66.1	-32.5
8.6	-27.1	77.0	-33.4
11.6	-27.5	79.6	-33.4
18.4	-28.3	82.4	-34.4
22.0	-28.4	87.0	-34.7
27.3	-28.6	90.4	-34.8
31.9	-30.5	94.0	-35.2
39.6	-31.2	95.9	-35.4
45.8	-31.3	98.9	-35.8
50.8	-31.6	100.0	-36.2

mol %	m.t.	mol %	m.t.
5.3	-27.9	56.3	-35.4
10.1	-29.4	61.8	-35.4
15.8	-30.8	71.2	-35.6
25.3	-33.4	81.0	-36.1
36.1	-34.8	90.4	-36.4
48.5	-35.4	96.0	-35.9

Toropov, 1956

mol %	d 20°	d 40°	d 60°	η 20°	η 40°	η 60°
0	1.7310	1.7014	1.6605	827	702	589
20	1.8340	1.8012	1.7574	847	702	591
40	1.9351	1.8990	1.8527	870	703	595
60	2.0326	1.9933	1.9444	895	705	604
80	2.1276	2.0849	2.0336	922	709	616
100	2.2201	2.1740	2.1202	951	715	635

Titanum chloride ($TiCl_4$) + Antimonium pentachloride ($SbCl_5$)

Nasu, 1933

mol%	f.t. (cooling)	E (cooling)	mol%	f.t. (heating)	E (heating)
0.0	-24.8	-	4.9	-27.3	-50.0
7.9	-28.7	-	18.3	-35.8	-50.2
14,7	-33.8	-	56.0	-34.0	-49.8
20.5	-36.8	-	63.3	-25.1	-50.2
25.6	-40.4	-	71.7	-18.1	-50.0
30.1	-44.4	-	78.2	-13.1	-50.0
34.1	-47.2	-	86.0	-8.0	-50.2
36.3	-48.7	-	93.5	-3.0	-50.0
37.6	-	-			
40.8	-45.7	-			
46.3	-41.4	-			
51.9	-35.4	-			
59.0	-29.4	-			
68.3	-21.3	-49.7			
74.2	-17.3	-50.2			
81.2	-11.3	-51.9			
100.0	+20	-51.7			

Titanium chloride ($TiCl_4$) + Tantalum chloride ($TaCl_5$)

Tarasenkov and Komandine, 1940

% L	% V	t	p	p_2
6.20	0.41	138.0	750	1.63
17.80	1.09	138.2	750	4.36
23.73	1.49	139.0	763	6.07
39.30	2.78	140.8	764	11.04

Tantalum chloride ($TaCl_5$) + Niobium chloride ($NbCl_5$)

Schäfer and Pietruck, 1952

mol%	f.t.	mol%	f.t.
0	216.5	60	209.9
10	215.5	70	208.6
20	214.5	80	207.1
30	213.4	90	205.8
40	212.3	100	204.7
50	211.1		

Morozov, 1956 (fig.)

%	f.t.	E	%	f.t.	E
100	203	-	40	208	204
80	205	202	20	212	205
60	207	203	0	220	-

Zirconium chloride ($ZrCl_4$) + Niobium pentachloride ($NbCl_5$)

Morozov, Korshunov, 1956.

t	p_1	p_2
	75.0%	
155	0.0	19.2
160	0.0	24.9
170	0.0	47.8
170.5	0.0	46.6
205	0.9	191.5
	50.0%	
155.5	0.0	20.8
160.5	0.0	23.9
170	0.0	42.4
170.5	0.0	44.2
205.5	2.3	200.0
	25.0%	
155.5	0.0	19.6
160	0.0	24.2
170.5	0.0	36.5
206	-	191.0

%	f.t.	E
100.0	207	-
90.0	195	186
80.0	193	186
70.0	-	183
60.0	285	186
40.0	388	185
30.0	405	187
20.0	416	187
10.0	433	182
0.0	437	-

XLI OTHER BINARY COMPOUNDS

Lithium bromide (LiBr) + Sodium bromide (NaBr)

Kellner, 1917

mol%	wt%	f.t.	m.t.
0	0	552	-
10	11.63	537	531
20	22.85	525	523
30	33.68	537	527
50	54.23	592	542
70	73.44	672	631
85	87.04	712	688
100	100	742	-

Lithium bromide (LiBr) + Potassium bromide (KBr)

Kellner, 1917

mol%	wt%	f.t.	E	min.
0	0	552	-	-
5	6.73	534	335	60
20	25.51	465	345	270
30	36.99	409	348	400
37	44.59	364	"	520
40	47.74	-	"	560
45	52.85	402	346	470
60	67.27	538	345	370
75	80.43	638	346	220
95	96.30	715	345	60
100	100	730	-	-

Lithium bromide (LiBr) + Rubidium bromide (RbBr)

Gromakov and Gromakova, 1953

mol%	f.t.	E	
0	527	-	
10	494	-	
20	432	259	
25	400	259	
30	364	259	
35	318	259	
40	265	259	
(41)	259	-	
42,5	269	259	E (1+1)
45	273	271	
50	335	271	
55	398	271	
60	450	271	
72.6	545	-	
89.6	642	-	
100	681	-	

Zirconium chloride ($ZrCl_4$) + Niobium pentachloride ($NbCl_5$)

Morozov, Korshunov, 1956.

t	p_1	p_2
	75.0%	
155	0.0	19.2
160	0.0	24.9
170	0.0	47.8
170.5	0.0	46.6
205	0.9	191.5
	50.0%	
155.5	0.0	20.8
160.5	0.0	23.9
170	0.0	42.4
170.5	0.0	44.2
205.5	2.3	200.0
	25.0%	
155.5	0.0	19.6
160	0.0	24.2
170.5	0.0	36.5
206	-	191.0

%	f.t.	E
100.0	207	-
90.0	195	186
80.0	193	186
70.0	-	183
60.0	285	186
40.0	388	185
30.0	405	187
20.0	416	187
10.0	433	182
0.0	437	-

XLI OTHER BINARY COMPOUNDS

Lithium bromide ($LiBr$) + Sodium bromide ($NaBr$)

Kellner, 1917

mol%	wt%	f.t.	m.t.
0	0	552	-
10	11.63	537	531
20	22.85	525	523
30	33.68	537	527
50	54.23	592	542
70	73.44	672	631
85	87.04	712	688
100	100	742	-

Lithium bromide ($LiBr$) + Potassium bromide (KBr)

Kellner, 1917

mol%	wt%	f.t.	E	min.
0	0	552	-	-
5	6.73	534	335	60
20	25.51	465	345	270
30	36.92	409	348	400
37	44.59	364	"	520
40	47.74	-	"	560
45	52.85	402	346	470
60	67.27	538	345	370
75	80.43	638	346	220
95	96.30	715	345	60
100	100	730	-	-

Lithium bromide ($LiBr$) + Rubidium bromide ($RbBr$)

Gromakov and Gromakova, 1953

mol%	f.t.	E	
0	527	-	
10	494	-	
20	432	259	
25	400	259	
30	364	259	
35	318	259	
40	265	259	
(41)	259	-	
42.5	269	259	E (1+1)
45	273	271	
50	335	271	
55	398	271	
60	450	271	
72.6	545	-	
89.6	642	-	
100	681	-	

Lithium bromide (LiBr) + Silver bromide (AgBr)

Sandonnini and Scarpa, 1913

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0.0	556	-	60.0	448	435
10.0	525	522	70.0	443	-
20.0	515	485	44.0	432	-
30.0	484	464	90.0	427	-
40.0	470	450	100.0	419	-
50.0	460	435			

Salstrom and Hildebrand, 1930

t	e	t	e
100 mol%		59.37 mol%	
442.3	0.8031	472.5	0.8153
453.6	.800	501.8	.8081
456.6	.7989	504.2	.8072
467.0	.7956	529.8	.8011
490.9	.7887	567.7	.7918
499.9	.7866	571.9	.7907
521.4	.7803	40.86 mol%	
524.4	.7795	503.8	0.8205
531.7	.7769	524.2	.8154
538.3	.7751	547.5	.8105
556.2	.7702	577.2	.8045
565.0	.7680	579.4	.8040
25.48 mol%		11.00 mol%	
514.9	0.8286	562.0	0.8639
546.3	.8220	566.3	.8635
549.7	.8213	589.5	.8616
575.3	.8171	612.8	.8603
596.0	.8137		

Lithium bromide (LiBr) + Magnesium bromide ($MgBr_2$)

Kellner, 1917

mol%	wt%	f.t.	m.t.	E	min.
0	0	552	-	-	-
10	19.07	551	548	-	-
20	34.64	543	515	-	-
30	47.61	537	511	-	-
40	58.57	542	532	-	-
45	63.43	552	-	548	160
50	67.95	574	-	"	130
55	72.15	594	-	"	100
60	76.08	620	-	"	90
70	83.18	646	-	"	60
80	89.45	675	678	-	-
90	95.02	694	-	-	-
100	100	711	-	-	-

Ferrari and Colla, 1931

mol%	f.t.	mol%	f.t.
100	692	25.0	537
76.3	640	17.7	540
52.9	580	11.2	548
44.9	530	5.3	549
42.9	528	0	550
33.4	533		

Lithium bromide (LiBr) + Calcium bromide ($CaBr_2$)

Kellner, 1917

mol%	wt%	f.t.	m.t.	E	min.
0	0	552	-	-	-
10	20.37	551	542	-	-
20	36.53	550	529	-	-
30	49.66	545	531	-	-
35	55.35	548	534	-	-
40	60.54	552	543	-	-
42.5	62.98	563	-	-	-
45	65.32	573	-	563	190
50	69.71	579	-	562	180
60	77.54	587	-	562	120
70	84.30	625	-	559	60
75	87.35	685	-	563	-
80	90.20	692	-	568	-
90	95.40	705	683	-	-
100	100	730	-	-	-

Lithium bromide (LiBr) + Strontium bromide ($SrBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.	tr.t.	min.
0	0	552	-	-	-	-
5	13.04	540	449	70	-	-
20	41.60	499	454	230	-	-
30	54.98	463	453	370	-	-
40	65.51	471	454	300	-	-
50	74.02	495	453	130	-	-
60	81.04	500	447	60	503	220
66.66	85.07	530	-	-	501	180
75	89.53	553	-	-	502	90
85	94.17	592	-	-	-	-
100	100	643	-	-	-	-

Lithium bromide (LiBr) + Barium bromide ($BaBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.
0	0	552	-	-
5	15.26	539	476	60
20	46.10	497	484	270
25	53.28	-	483	370
40	69.52	563	484	320
60	83.69	672	482	230
80	93.19	780	478	90
95	98.49	825	474	40
100	100	847	-	-

Lithium bromide (LiBr) + Lead bromide ($PbBr_2$)

Gromakov, 1950

mol%	f.t.	E	mol%	f.t.	E
100	375	-	42.0	435	322
85.5	345	-	30.7	462	-
79	322	-	23.6	481	-
74.2	335	322	20	571	552
65.0	359	322	10	589	577
55.7	389	322	0	545	545
(1+1)					

Lithium bromide (LiBr) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	62.1	152.2
99.38	107.2 (1+7)	60.9	157.9
98.0	108.2	59.3	164.5
95.7	109.8	55.1	180.7
91.3	112.4	53.2	186.7
88.9	113.6	51.6	192.5
86.0	114.6	49.4	195.4
83.8	113.0	48.6	221.4
82.9	117.7 (1+2)	43.3	405.0
78.0	121.9	36.1	510.0
72.5	125.2	25.4	523.0
72.2	126.7	10.0	432.0
71.3	127.9	0	535.0
69.4	129.4		
65.8	135.5 (1+1)		

Sodium bromide (NaBr) + Potassium bromide (KBr)

Kurnakov and Zhemchuzhni, 1907

mol%	f.t.	mol%	f.t.
100.00	757	43.27	649
87.16	721	35.53	661
77.23	695	25.33	686
67.95	672	11.28	734
58.56	651	0.00	768
50.00	644		

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ		mol%	σ	
	700°	800°		700°	800°
0	109	103	60	96	89
20	103	95	80	95	86
40	99	92	100	94.5	85

Zhemchuzhni and Rambach, 1910

mol%	Q mixed cryst.
12	480
25	705
50	712
75	650
88	400
after 2 1/2 months	
50	250

Sodium bromide (NaBr) + Rubidium bromide (RbBr)

Gromakov and Gromakova, 1953

mol%	f.t.	E	f.t.
100	679	679	A
90	648	-	"
85	628	-	"
80	613	-	"
75	589	-	"
70	567	-	"
65	546	-	"
60	526	495	
55	502	495	
53.5	495	-	A + B
50	513	495	B
45	540	495	"
40	566	495	"
35	590	-	"
30	615	-	"
25	646	-	"
20	668	-	"
10	715	-	"
0	754	-	"

Sodium bromide (NaBr) + Silver bromide (AgBr)

Sandonnini and Scarpa, 1913

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0.00	748	-	60.0	515	430
10.0	720	-	70.0	490	425
20.0	672	-	80.0	440	425
30.0	630	-	90.0	420	410
40.0	600	-	100	419	-
50.0	560	430			

Zhemchuzhni, 1916

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	419	-	48	602	442
95	426	420	40	634	453
90	437	420	37	640	460
85	447	422	30	667	500
80	470	426	20	699	550
75	490	426	15	712	603
70	517	430	10	732	641
60	561	432	5	756	700
55	582	434	0	766	

Tubandt and Aramowitsch, 1927

mol%	f.t.	300°	350°	400°
100	420	17	70	-
90	460	10	40	-
80	510	6	24	68
70	540	3	13	38
60	580	2	7	20
50	620	0.5	5	10
40	650	0	2	6
30	690	-	0.5	2
20	710	-	-	-
10	740	-	-	-
0	765	-	-	-

Kolotii, 1956.

mol%	780°	e 790°	800°
0.23	0.4794	0.4844	0.4909
1.60	0.3501	0.3543	0.3582
4.76	0.2546	0.2548	0.2573
33.18	0.0495	0.0504	0.0515

mol%	810°	e 820°	830°
0.23	0.4936	0.5032	0.5090
1.60	0.3637	0.3666	0.3708
4.76	0.2602	0.2635	0.2663
33.18	0.0.0525	0.0532	0.0547

Sodium bromide (NaBr) + Magnesium bromide ($MgBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.
0	-	742	-	-
10	16	702	431	120
20	30.91	656	432	310
30	43.40	567	432	380
40	54.40	438	429	470
50	64.15	470	431	520
60	72.86	565	431	390
70	80.68	629	430	286
80	87.74	662	428	240
90	94.15	690	428	110
100	-	711	-	-

Sodium bromide (NaBr) + Calcium bromide ($CaBr_2$)

Kellner, 1917

%	f.t.	m.t.	min.	tr.t.	min.
0	742	-	-	-	-
17.75	717	700	-	-	-
32.69	677	646	-	-	-
45.43	631	E	-	-	-
51.12	613	500	-	-	-
56.43	584	499	40	442	40
61.38	555	505	100	466	"
66.02	534	514	170	466	"
70.36	523	513	180	469	"
74.45	-	514	200	468	"
78.30	563	511	130	469	50
79.53	564	509	100	470	60
81.93	609	513	100	469	30
88.60	650	508	70	468	30
94.59	694	-	-	-	-
100	730	-	-	-	-

Sodium bromide (NaBr) + Strontium bromide ($SrBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.
0	0	742	-	-
5	11.24	727	477	50
15	29.79	695	486	130
30	50.75	637	486	220
45	66.31	567	485	330
61	79.00	-	486	450
75	87.83	546	486	260
90	95.58	607	482	90
100	100	643	-	-

Sodium bromide (NaBr) + Barium bromide ($BaBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.
0	0	742	-	-
5	13.19	726	591	65
15	33.76	688	599	190
30	53.31	632	598	380
45	70.26	626	600	440
60	81.25	692	599	290
80	92.03	777	600	120
95	98.21	831	600	40
100	100	847	-	-

Sodium bromide (NaBr) + Lead bromide ($PbBr_2$)

Gromakov, 1950

mol %	f.t.	E	mol %	f.t.
100	375	-	55	471
90	349	-	50	503
85	336	306	45	529
80	318	-	40	566
78	306	306	35	591
75	333	306	30	609
70	378	-	25	633
65	413	-	20	655
60	444	-	0	745
		(1+1)		

Sodium bromide (NaBr) + Cadmium bromide ($CdBr_2$)

Brand, 1913

%	f.t.	E	min.	%	f.t.	E	min.
100	567	-	-	73.81	449	368	430
95.97	543	367	90	53.17	552	367	330
91.36	520	367	180	39.61	633	368	220
86.06	490	367	250	22.71	694	368	150
79.87	448	367	340	0	746	-	-
72.56	387	367	460				

Sodium bromide (NaBr) + Mercuric bromide ($HgBr_2$)

Belyaev and Mironov, 1952

mol %	f.t.	mol %	f.t.	mol %	f.t.
100	243.0	90.0	238.0	81.2	281.5
95	239.0	88.0	239.0	80.0	286.0
93	236.0	86.0	242.5	76.5	306.0
91.5	232.5	84.0	256.5	75.0	321.0
90.5	235.0	82.0	277.5	74.2	325.5

E : 91.5 mol % 232° tr.t.$_1$: 85 mol % 243°
tr.t.$_2$: 79 mol % 288°

Sodium bromide (NaBr) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.		mol%	f.t.	
100	97.1		81.6	94.0	
99.3	95.2		79.5	95.4	
99.1	94.8		78.5	95.8	
98.6	93.6		77.9	96.0	
98.3	93.0	compl.	76.6	95.1	
98.1	93.9		76.0	94.8	(1+2)
97.8	94.6		74.9	98.2	
96.9	110.9		72.5	102.4	
96.0	166.4		68.1	104.5	
94.6	227.1		67.4	108.8	(1+1)
92.2	231.9	L_1+L_2	64.7	131.2	
89.6	230.5		63.2	141.5	
87.2	202.7		61.9	154.8	
85.5	166.9		58.0	170.4	
84.1	125.6		55.0	184.0	
83.0	94.2		51.4	196.4	
82.3	93.0		49.7	200.5	
81.6	91.6		49.7	209.0	
81.0	90.4		48.9	360.0	
83.1	92.8	(2+7)			

97.4 mol % - 83.7 mol % 95.4° $L_1 + L_2 + C$

C.S.T.: 92.0 mol % 232°

Gorenbein and Kriss, 1949

%	d 120°	d 130°	d 140°
95.85	-	2.770	-
88.85	2.786	-	2.756
87.90	2.798	2.784	2.767
86.82	2.809	2.793	-
85.59	2.817	2.803	2.787
84.71	2.820	2.806	2.792
83.89	-	-	2.797
83.83	2.827	2.812	-

Gorenbein and Kriss, 1949

%	η 120°	η 130°	η 140°
95.85	-	11030	-
88.85	13340	-	9290
87.90	13730	11320	9690
86.82	14100	-	-
85.59	14760	12000	9940
84.71	15100	-	-
83.89	-	-	10330
83.83	15320	12350	-

Gorenbein and Kriss, 1949

%	κ 120°	κ 130°	κ 140°
95.85	-	401.6	-
88.85	345.9	-	456.8
87.90	382.3	444.4	500.0
86.82	441.3	505.7	-
85.59	491.7	565.0	641.4
84.71	533.4	613.3	686.7
83.89	-	-	778.3
83.83	577.6	673.9	-

Gorenbein and Kriss, 1949

%	molar conductivity 120°	130°	140°
95.85	-	13.38	-
88.85	11.46	-	15.29
87.90	11.59	13.57	15.36
86.82	12.23	14.16	-
85.59	12.30	14.20	16.21
84.71	12.74	14.72	16.56
83.89	-	-	17.70
83.83	13.00	15.26	-

Potassium bromide (KBr) + Rubidium bromide (RbBr)

Gromakov and Gromakova, 1953

mol%	f.t.	mol%	f.t.
100	680	50	685
90	675	40	691
80	668	30	700
75	670	20	708
70	673	(10)	719
60	678	0	730

Potassium bromide (KBr) + Cuprous bromide (CuBr)

De Cesaris, 1911

%	f.t.	tr.t.	min.	E	min.
0	730	-	-	-	-
5	718	232	120	182	120
10	706	232	150	182	180
20	662	232	195	182	300
30	615	234	225	180	375
35	574	238	240	183	405
40	548	234	300	184	450
45	500	234	240	182	480
50	446	234	180	182	555
60	342	236	60	184	690
65	222	-	-	183	750
70	212	-	-	183	675
80	302	-	-	182	450
90	420	-	-	182	240
95	442	-	-	182	150

Potassium bromide (KBr) + Thallium bromide (TlBr)

Rostkovski, 1929

mol%	f.t.	mol%	f.t.
100	457	77	520
98.88	457	73.2	535
95.5	458	69.1	546
93.3	459	63.4	565
89.4	469	58.4	579
87	479	52.8	600
83.3	497	43.5	614

Potassium bromide (KBr) + Silver bromide (AgBr)

Zhemchuzhni, 1916

mol%	f.t.	E	mol%	f.t.	E
100	419	-	40	536	283
98	408	279	30	604	282
95	394	280	20	654	282
90	373	283	10	695	282
80	330	285	5	724	282
70	289	285	2	738	280
60	372	285	0	748	-
50	465	283			

Tubandt and Abramovitsh, 1927

mol%	f.t.	mol%	f.t.
100	420	50	410
90	395	40	490
80	340	30	570
70	295	20	670
65	-	10	690
63	-	0	760
60	340		

Lifshits, 1956.

mol%	f.t.
100	419
69	286 E
0	740

Tubandt and Abramovitsh, 1927

mol%	κ 200°	κ 230°	κ 250°	κ 270°
100	7.7	22	39	76
90	6.1	14	29	46
80	5.1	13	23	36
70	4.6	11	20	33
65	4.3	-	-	-
63	-	-	19.5	31
60	3.2	9	14	23
50	2.2	6	8	16
40	1.4	4	7	11
30	1.0	-	-	8
20	-	-	-	4
10	-	-	-	0
0	-	-	-	-

Harrap and Heyman, 1955

mol%	η	κ(in mhos)	η	κ(in mhos)
	350°		400°	
85.1	-	1.816	-	1.975
79.7	4440	-	3550	-
69.2	4640	1.256	3550	1.420
	450°		500°	
100	3260	2.904	2830	3.026
85.1	-	2.114	-	3.232
79.7	2900	-	2460	-
69.2	2860	1.572	2400	1.698
55.2	2860	-	2350	-
50.7	-	-	-	1.290
	550°		575°	
100	2510	3.124	2380	3.168
85.1	-	2.328	-	2.380
79.7	2120	-	1980	-
69.2	2050	1.802	1910	1.854
55.2	1980	-	1850	-
50.7	-	1.424	-	1.486
45.7	2040	-	1870	-
40.9	-	1.304	-	1.374
	600°			
100	2270	3.210		
85.1	-	2.428		
79.7	1860	-		
69.2	1780	1.906		
55.2	1740	-		
50.7	-	1.546		
45.7	1730	-		
40.9	-	1.428		

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ 700°	σ 750°	mol%	σ 700°	σ 750°
0	94.5	90	60	124	118.5
20	96	92.5	80	137.5	131.5
40	99	96	100	148	146

Potassium bromide (KBr) + Magnesium bromide ($MgBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.	tr.t.	min.
0	0	730	-	-	-	-
10	14.67	703	330	110	343	-
20	27.89	629	331	190	348	-
25	34.03	573	332	310	345	-
30	39.87	359	334	370	353	-
33.33	43.62	348	333	420	353	-
35	45.45	-	334	590	353	-
40	50.78	358	333	370	-	-
45	55.87	389	331	170	-	-
50	60.74	409	-	-	391	310
60	69.89	536	-	-	391	230
70	78.31	580	-	-	391	210
80	86.09	635	-	-	391	130
90	93.30	688	-	-	377	90
100	100	711	-	-	-	-

Potassium bromide (KBr) + Calcium bromide ($CaBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.	
0	0	730	-	-	
10	15.73	693	544	120	
20	29.58	648	545	260	
30	41.86	586	542	310	
40	52.83	619	543	170	
50	62.68	637	-	-	(1+1)
60	71.59	614	562	100	
70	79.67	580	567	190	
80	87.04	634	560	140	
90	93.80	689	561	80	
100	100	730	-	-	

Potassium bromide (KBr) + Strontium bromide ($SrBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.	
0	0	730	-	-	
7	13.59	709	556	90	
10	18.77	699	558	120	
25	40.94	599	556	220	
33.33	50.97	519	-	-	(1+1)
40	58.09	553	533	110	
50	67.52	-	534	330	
57	73.38	559	529	110	
66.67	80.61	574	-	-	
75	86.18	570	563	100	
85	92.18	578	561	160	
95	97.53	624	539	50	
100	100	643	-	-	

Potassium bromide (KBr) + Barium bromide ($BaBr_2$)

Kellner, 1917

mol%	wt%	f.t.	E	min.	
0	0	730	-	-	
5	11.62	712	629	70	
15	30.59	673	632	200	
20	38.43	646	634	320	
28	49.27	633	633	230	
33.33	55.53	634	-	-	
42.5	64.86	626	614	180	
50	71.41	612	611	330	
60	78.93	655	613	260	
70	85.35	712	608	190	(1+1)
80	90.90	770	614	120	
95	97.94	835	594	30	
100	100	847	-	-	

Potassium bromide (KBr) + Lead bromide ($PbBr_2$)

Gromakov, 1950

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	375	375	52.5	350	-
90	346	334	50	369	350
86	334	-	45	396	350
80	364	334	42.5	408	-
75	374	-	40	428	408
70	378	-	35	492	409
67.7	380	-	30	545	409
65	379	-	25	586	408
60	372	-	20	617	-
55	360	350	0	730	730

(2+1) and (1+2)

Pokotilo, N.M. and K.I., 1940

mol %	f.t.	E	mol %	f.t.	E
100	372	-	56.9	378	-
95.1	361	349	55.5	376	-
90.5	352	-	54.2	368	368
87.9	349	-	51.0	392	-
86.1	354	-	48.2	409	-
82.0	378	-	45.4	420	-
79.0	384	-	44.5	425	-
77.8	385	-	42.9	450	410
73.7	392	-	38.0	500	-
70.2	394	-	32.9	551	410
66.7	395	368	24.7	614	410
63.3	391	-	17.4	660	410
59.6	389	-	0	740	-

Lorenz, Frei and Jabs, 1908

d = 3.48-t/600 + (2.786+t/15000) - N/100 where
N = mol % lead bromide

Potassium bromide (KBr) + Cadmium bromide ($CdBr_2$)

Brand, 1913

%	f.t.	E	min.	tr.t.	min.
100	567	-	-	-	-
95.37	536	344	100	-	-
90.15	508	344	200	-	-
87.28	488	345	230	-	-
84.23	464	345	280	-	-
77.43	397	346	340	-	-
71.25	350	345	130	-	-
69.58	354	-	-	-	-
65.17	348	304	100	-	-
60.39	329	305	200	-	-
55.19	315	305	529	-	-
53.35	319	304	450	-	-
49.50	443	304	400	-	-
45.45	514	305	300	323	50
43.26	544	303	250	325	80
36.38	605	304	210	324	100
28.76	652	306	150	325	120
20.26	682	304	120	324	90
10.75	712	300	40	324	80
0	735	-	-	322	50
			(1+1)	-	-

Potassium bromide (KBr) + Mercuric bromide ($HgBr_2$)

Belyaev and Mironov, 1952

mol%	f.t.	tr.t.	mol%	f.t.	tr.t.
100	244.0	-	56.5	175.0	-
95.0	238.5	218	56.3	176.0	-
90.0	232.0	218	55.0	182.0	-
87.5	227.0	-	54.6	182.0	-
85.0	222.0	218	53.5	184.0	-
83.5	219.0	-	52.9	187.0	-
82.0	217.0	185	52.0	188.0	-
80.0	215.0	185	51.5	191.0	-
77.5	210.0	-	50.0	190.0	-
75.0	206.0	185(1+2)	49.0	192.0	-
72.5	203.0	-	47.5	201.0	-
70.0	199.0	-	46.0	206.5	-
68.5	193.0	-	45.0	214.0	190
65.0	183.0	-	44.0	235.0	-
62.5	182.0	168	42.5	288.0	-
60.0	176.0	168	40.0	362.0	207
58.2	168.0	-	35.0	476.0	207
58.0	171.0	168	30.0	561	-
57.5	171.0	168 E			

(1+1) and (1+2)

Potassium bromide (KBr) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t
100	97.1	65.3	109.0 (1+1)
99.67	97.5 compl.	62.7	130.6
99.14	265.9	60.5	145.9
80.6	189.6	59.0	154.6
76.7	95.6	55.5	171.4
75.3	92.2	52.6	181.7
74.2	88.6 (1+2)	51.0	188.4
72.9	90.8	48.7	189.6
71.2	93.0	48.0	188.8
68.4	95.5	45.3	390.0
66.8	95.8		

99.6-77.9mol% L_1+L_2 + C(complex) 98.1°

Izbekov and Plotnikov, 1911

%	d
99.5°	
88.07	2.803
80.99	2.830

Gorenbein, 1945

wt %	mol %	d	wt %	mol %	d
		100°			
100	100	2.644	83.86	69.85	2.825
88.80	77.80	2.798	81.76	66.66	2.830
85.50	72.40	2.824			

Gorenbein and Kriss, 1949

%	d 110°	120°	130°	140°
88.00	2.790	2.774	2.757	2.741
87.45	2.798	2.780	2.764	2.749
86.24	-	-	-	2.755
84.72	-	2.794	2.779	2.764
83.15	2.815	2.804	2.785	2.771
81.765	2.818	2.802	2.789	2.775

Gorenbein, 1945

wt %	mol %	η	wt %	mol %	η
		100°			
100	100	2394	83.86	69.85	25185
88.80	77.80	19040	81.76	66.66	25698
85.50	72.40	23924			

Gorenbein and Kriss, 1949

%	η 110°	120°	130°	140°
88.00	15840	12630	10490	8650
87.45	16620	13300	10870	9060
86.24	17820	14090	11760	9740
84.72	18960	14970	12260	10140
83.15	19190	15200	12370	10500
81.765	-	15730	12930	10820

Izbekow and Plotnikov, 1911 and Izbekow and Nijnik, 1937

%	ϰ	%	ϰ
	99.5°		
88.98	163.3	83.16	236.2
87.35	185.3	81.54	254.6
84.56	219.5		

Gorenbein and Kriss, 1949

%	ϰ 110°	120°	130°	140°
88.00	233.1	278.1	324.3	378.5
87.45	-	290.3	345.3	-
86.24	268.7	322.4	378.2	438.9
84.72	300.9	-	428.1	492.8
83.15	330.8	392.2	473.5	545.9
81.765	360.3	431.5	513.8	600.9

Gorenbein and Kriss, 1949

%	molar conductivity 110°	120°	130°	140°
88.00	8.29	9.94	11.67	13.70
87.45	-	9.91	11.85	-
86.24	8.29	9.97	11.80	13.78
84.72	8.34	-	12.00	13.89
83.15	8.30	10.04	12.01	13.91
81.765	8.34	10.05	12.02	14.13

Rubidium bromide (RbBr) + Lead bromide ($PbBr_2$)

Gromakov, 1950

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	375	375	50	400	400
90	338	325	47	396	-
85	338	325	45	388	-
80	355	-	40	374	374
75	364	-	38	405	-
70	368	-	35	425	-
66.7	370		30	478	432
65	368	348	25	536	433
60	366	348	20	574	-
55	390	348	0	680	680

(1+1) (1+2) (2+1)

Thallium bromide (TlBr) + Lead bromide ($PbBr_2$)

Favorskii, 1941

mol%	f.t.	mol%	f.t.
100	373	48.2	366
95.0	363	47.1	360
90.2	350	46.0	358
85.2	366	45.1	360
83.2	371	44.1	364
81.4	376	42.0	371
79.5	382	40.0	375
77.0	385	38.0	379
76.4	386	35.0	384
75.2	388	33.4	386
74.0	391	31.0	387
73.4	392	29.1	390
71.0	393	27.0	392
69.2	394	25.9	392
66.7	395	25.0	392
65.1	394	24.0	392
62.2	392	23.0	390
60.0	391	20.9	387
58.2	386	19.0	388
56.1	384	17.0	400
54.0	380	14.9	408
52.1	375	14.0	417
50.1	371	9.8	427
49.1	368	5.1	443 (1+2)

Thallium bromide (TlBr) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	62.4	142.1
99.94	260.0 compl.	60.0	157.7
78.6	260.0	59.3	160.4
77.6	118.4	57.8	171.8
75.6	99.9	55.7	192.9
74.0	104.4 (1+2)	51.8	203.4
73.1	105.9	49.0	207.9
71.6	106.7	47.9	200.8
69.3	108.1	46.7	199.1
67.3	110.6	45.0	213.8
64.6	111.8	44.8	215.9
64.6	126.8 (1+1)		

99.4 - 77.2mol% L_1+L_2+C (complex) 103.9°

Cuprous bromide (CuBr) + Cadmium bromide ($CdBr_2$)

Herrmann, 1911

%	f.t.	m.t.	E	min.	tr.t.	min.
0	484	-	-	-	394	8
10	470	461	-	-	371	9
30	437	430	-	-	353	9
50	428	422	-	-	349	9
70	441	-	-	-	351	9
72.5	462	-	-	-	350	10
75	467	-	422	7.5	349	8
80	494	-	420	6	345	6
90	532	-	420	3	344	4
100	568	-	-	-	-	-

Silver bromide (AgBr) + Lead bromide ($PbBr_2$)

Tubandt and Eggert, 1920

mol%	f.t.	E	f.t. (1+2)
0	422	-	-
10	395	274.8	-
20	366	275.5	-
30	330	275	-
40	296	276.3	-
50	280	275.4	-
60	290	275.2	-
66.67	295	(275.2)	-
70	300	-	295
80	317	-	295
90	344	-	295
100	373	-	-

(1+2)

Silver bromide (AgBr) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	74.4	115.6
98.9	94.2	73.6	117.8
98.7	93.4	69.8	125.2 (1+1)
98.5	95.3 (1+2)	68.0	135.0
98.4	98.2	64.4	154.6
97.0	120.4 L_1+L_2	62.7	164.1
95.5	156.3	60.8	174.1
95.0	161.9	59.1	182.3
92.4	183.6	57.3	189.5
86.6	173.0	53.2	206.8
85.3	159.7	49.8	215.8
83.8	139.7	49.4	214.2
81.3	108.6	48.0	210.9
79.3	111.3	45.1	219.0
77.9	112.8		

97.8 mol% - 83.0 mol% 105.9° L_1+L_2+C

C.S.T.: 90.2 mol% 186°

Mercurous bromide (HgBr) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	55.6	255.1
99.4	696.6	51.4	259.7
98.8	161.5	46.8	256.6
98.3 L_1+L_2	225.4	45.3	252.9
96.7	275.5	40.4	241.4
66.1	242.1	37.3	243.7
65.0	243.9	33.9	281.3
59.5	250.1		

98.2 mol% - 69.2 mol% L_1+L_2+(1+1) 238.1°

Magnesium bromide ($MgBr_2$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	91.9	199.6
99.4	96.8	87.1	210.5
98.6	96.5	81.6	221.6
99.4	134.9 (1+2)	78.4	227.9
98.6	160.7	76.5	235.5
96.2	190.2		

Calcium bromide ($CaBr_2$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	78.6	242.6
99.26	204.2 compl.	75.7	260.1
84.6	195.3	72.7	276.8
84.0	213.1 (1+2)	69.1	298.4
81.1	229.5	66.2	304.9

(1+2) 99.2 mol% - 86.0 mol% L_1+L_2 208.8°

Barium bromide ($BaBr_2$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.
100	97.1
81.7	276.7 (1+2)
78.8	292.0
75.8	310.0
72.0	335

99.1 - 84.0 $L_1 + L_2$ + C(1+2) 269.4°

Zinc bromide ($ZnBr_2$) + Cadmium bromide ($CdBr_2$)

Zakharchenko, 1951

%	f.t.	E	%	f.t.	E
0.00	395	-	51.49	457	364
5.92	390	-	64.07	495	364
17.13	377	364	79.00	511	-
27.50	376	"	100.00	575	-
34.57	402	"			

Zinc bromide (ZnBr) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	87.9	96.1
97.4	95.5	85.5	99.3
94.4	94.4	79.4	104.1
88.7	87.6	73.3	108.8
85.5	83.5	70.0	110.4
88.7	95.1	64.9	110.6

Gorenbein, 1945

mol%	η 150°	140°	130°	120°	110°
100	1000	1200	1400	1600	1800
80	5000	6000	8000	10000	13000
75	10000	12000	15000	20000	30000
73	13000	18000	23000	30000	40000
68	24000	30000	48000	75000	-

mol%	κ 100°	110°	120°	130°	140°	150°
88	0.5	0.5	0.5	0.5	0.5	0.5
80	3	5	7	8	9	10
70	6	7.5	9.5	12.5	16	19

wt%	mol%	d 100°	110°	120°
70.46	66.80	3.009	2.993	2.971
74.49	71.14	-	.936	.917
77.75	74.68	2.914	.891	.872
80.22	77.38	.881	.859	.839
82.52	79.94	.849	.827	.806
84.97	82.67	.811	.791	.770
88.39	86.54	.768	.744	.723
100	100	.644	.625	.601

wt%	mol%	d 130°	140°	150°
70.46	66.80	2.952	2.934	2.915
74.49	71.14	.901	.875	.850
77.75	74.68	.853	.830	.811
80.22	77.38	.818	.796	.750
82.52	79.94	.785	.762	.740
84.97	82.67	.751	.726	.704
88.39	86.54	.702	.679	.657
100	100	.578	.556	.524

Cadmium bromide ($CdBr_2$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	86.0	197.0
99.3	95.6	80.7	204.2
98.3	140.7 (1+2)	74.0	217.4
98.2	168.1	71.1	221.2
95.5	189.6	66.9	224.0
94.8	190.0	64.7	223.1
91.0	192.9	64.7	234.9
88.4	195.4		

Mercuric bromide ($HgBr_2$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol %	f.t.	mol %	f.t.
	(1+2) I		
100	97.1	74.2	102.8
98.6	95.8	71.3	103.6
96.2	94.3	68.6	104.1
92.9	95.9	67.2	103.9
89.9	96.7	62.3	103.1
86.3	98.7	59.7	101.9
82.1	100.1		
	(1+2) II		
95.4	93.5	54.8	123.2
92.9	94.7	50.6	145.2
90.2	95.8	40.4	175.0
82.1	99.1	37.2	183.6
79.9	100.1	24.3	206.8
74.2	101.9	16.9	217.1
71.3	102.5	12	224.4
67.8	102.6	6	232.9
59.7	100.7	1.9	239.9
55.8	118.8	0	241.5

Izbekow, 1925 and Nijnik, 1937

mol%	f.t.	mol%	f.t.
100	97.4	55.3	120.4
95.5	94.2	49.2	142.2
93.9	93.5	44.0	168.9
90.2	93.9	36.4	184.0
85.1	96.5	30.1	197.1
80.8	97.8	23.3	209.4
76.0	100.3	15.6	220.2
71.4	102.5	9.3	227.3
66.7	102.8	0	241.6
61.8	100.5		
		(1+2)	

Izbekov and Plotnikov, 1911

%	d
99.5°	
90.04	2.827
81.58	3.028
71.86	3.263

%	κ	%	κ
	99.5°		
98.62	0.001	90.72	3.299
96.94	.004	87.59	12.165
94.48	.088	83.96	28.465
91.86	1.489	75.87	78.775
89.52	6.017	70.20	113.900
85.74	19.780	92.19	1.208
81.66	41.680	85.62	20.510
78.89	58.380	78.92	58.14
72.85	97.990		

Gorenbein, 1947

%	d 110°	120°	130°	140°
59.68	3.577	3.553	3.532	3.504
62.78	.502	.474	.451	.428
65.40	.415	.390	.370	.359
67.28	.368	.344	.325	.298
71.17	.276	.253	.228	.202
74.92	.173	.149	.124	.097
80.50	.030	.010	2.982	2.961
89.56	2.827	2.800	.779	.755
100	.624	.599	.578	.555

%	η 110°	120°	130°	140°
59.68	25197.7	18321.5	13785.9	10886.4
62.78	22357.4	16599.6	13120.7	10316.9
65.40	19579.1	14678.2	11492.0	9172.0
67.28	17836.0	13905.0	10928.0	8546.0
71.17	14328.3	11125.1	8762.7	6992.8
74.92	12183.0	9146.0	7192.0	5827.0
80.50	7349.0	6066.0	4801.0	4108.0
89.56	3726.7	3165.5	2707.2	2319.6
100	2168.0	1946.1	1762.1	1610.0

%	κ 110°	120°	130°	140°
59.68	138.40	169.747	211.805	245.364
60.51	133.947	167.446	202.458	234.222
63.92	121.903	146.680	174.121	200.384
67.96	100.479	120.275	140.957	161.317
79.01	41.938	48.165	52.896	57.278
89.56	6.421	6.702	7.145	6.414

Mercuric bromide ($HgBr_2$) + Lead bromide ($PbBr_2$)

Sandonnini, 1911

mol%	f.t.	E	min.
0	238	-	-
2.5	236	232	40
5	E	232	200
6	240	233	120
7.5	244	234	100
10	252	232	80
20	271	233	70
30.4	279	233	80
41.2	283	232	60
51.1	289	232	50
63.0	301	232	40
73.3	314	232	30
83.8	322	232	20
100	366	-	-

Mercuric bromide ($HgBr_2$) + Antimonium tribromide ($SbBr_3$)

Nijnik, 1937

mol%	f.t.	mol%	f.t.
100	97.4	100	97.4
95.5	94.2	94.5	94.2
93.8	93.5	90.1	93.2
98.2	93.9	84.1	90.6
85.1	96.5	75.5	84.4
84.8	97.8	71.4	80.0
76.0	100.3	69.0	74.7
71.4	102.5	65.2	75.0
66.5	102.8	62.4	76.6
61.8	100.5	48.6	82.2
55.3	120.4	42.0	79.8
49.2	142.2	35.2	74.6
44.0	168.9	27.5	71.0
36.4	184.0	23.0	75.8
30.2	197.1	15.4	85.4
23.3	209.4	8.4	90.5
15.6	220.2	0	95.5
9.3	227.3		

Stannous bromide ($SnBr_2$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

%	f.t.	%	f.t.	
100	97.1	71.7	190.2	
99.65	96.9	69.3	197.8	
99.19	96.5	65.1	202.0	
98.93	96.3	58.5	181.8	
99.19	121.6	56.3	175.4	
98.93	137.1	55.2	175.0	(1+1)
98.55	152.8	51.7	179.8	
97.9	169.1	47.6	179.1	(1+1)
96.6	187.3	44.8	172.9	
95.3	198.0	40.4	164.4	
89.6	202.4	37.2	158.3	
89.3	201.4	28.8	175.1	
86.7	185.4	21.9	195.9	(1+2)
83.6	162.2	17.1	206.4	
83.4	162.7	9.2	220.9	
81.5	164.3	8.8	228.2	
75.3	178.5	0	232.0	

$L_1 + L_2$ + complex 98.2 - 85.8 mol % 161.1°

C.S.T. : 91.6 mol % 204.5°

Lead bromide ($PbBr_2$) + Cadmium bromide ($CdBr_2$)

Zakharchenko, 1951

%	f.t.	E	%	f.t.	E
100	375	-	53.10	486	344
97.26	371	-	39.87	512	"
84.50	356	344	21.41	542	-
78.20	360	"	11:02	536	-
71.50	426	"			

Lead bromide ($PbBr_2$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
0.63	191.9	37.0	266.9
16.9	211.9	39.0	257.0
20.1	220.4	43.6	241.6
23.7	235.5	45.5	234.9
27.0	253.5	52.3	268.4
30.5	267.7	57.7	296.8
32.5	272.5		

(1+2)+ L_1+L_2 210.4° 99.2 - 83.8mol%

Lead bromide ($PbBr_2$) + Bismuth bromide ($BiBr_3$)

Herrmann, 1911

%	f.t.	E	min.
0	380	-	-
10	373	-	-
20	337	-	-
25	324.5	205	0.5
30	312	201	1.4
50	271	207	5
60	252	207	7
70	218	205.5	9
80	-	206	12
90	210	205	5
100	219	-	-

Manganese bromide ($MnBr_2$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.		mol%	f.t.
100	97.1		82.5	217.6
99.32	96.3		79.4	223.8
99.32	127.1	(1+2)	76.0	232.9
98.02	171.0		74.2	237.7
95.4	199.1		72.0	241.7
94.4	199.8		70.4	242.6
90.6	201.6		69.0	300.0
86.2	210.8			

Aluminum bromide ($AlBr_3$) + Antimonium tribromide ($SbBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
0	97.1	63.5	78.8
8.2	94.7	66.4	76.9
15.2	92.3	72.5	72.9
19.2	90.3	75.7	76.3
28.6	82.9	81.68	81.9
37.1	80.5	86.7	86.5
46.5	84.2	92.8	91.3
49.3	84.1	100.0	96.6
53.5	84.3		(1+1)

Izbekov, 1925 and Nijnik, 1937

mol%	f.t.	mol%	f.t.
0	97.4	51.4	82.2
5.5	94.2	58.0	79.8
9.9	93.2	64.8	74.6
15.8	90.6	72.5	71.0
24.5	84.4	77.0	75.8
28.6	80.0	84.6	85.4
31.0	74.7	91.6	90.5
34.8	75.0	100	95.5
37.6	76.6		

(1+1)

Leggett, Vold and Mc Bain, 1942

Clearing point curve .

Gorenbein, 1945

mol%	η 140°	130°	120°	110°	100°
0	1500	1800	2000	2100	2200
10	2000	2500	2800	3200	3500
25	4500	5500	6500	8000	12000
46.5	6000	7500	9000	12000	15500
75	4000	4500	5500	7000	8000
100	2000	2300	2600	3200	3600

mol%	κ 100°	110°	120°	130°	140°
10	10	8	5	2	0
11	15	15	15	15	15
25	50	60	80	90	100
50	120	140	155	180	205
60	160	190	210	235	260
65	180	200	220	240	255
70	185	190	210	235	250
80	170	175	180	190	200

wt%	mol%	d 100	110	120	130	140
100	100	3.697	3.673	3.644	3.619	3.594
96.57	95.61	.694	.645	.620	.596	.572
83.10	84.53	.607	.586	.561	.537	.508
80.52	75.31	.543	.521	.500	.472	.449
77.26	71.50	.570	.486	.462	.442	.416
72.99	65.94	.473	.453	.430	.407	.385
67.36	60.37	.416	.395	.372	.348	.326
65.92	58.70	.402	.379	.356	.334	.311
57.54	49.92	.318	.298	.274	.249	.224
53.96	46.37	.279	.259	.236	.214	.190
52.55	44.97	.267	.246	.224	.203	.182
45.76	38.36	.193	.171	.146	.123	.099
41.48	34.35	.136	.115	.091	.068	.045
32.60	26.31	.034	.013	2.990	2.966	2.941
28.05	22.34	2.971	2.950	.928	.905	.881
25.74	20.38	.940	.918	.896	.871	.847
22.01	17.23	.898	.877	.853	.882	.804
20.90	16.31	.881	.859	.837	.815	.793
16.42	12.89	.833	.811	.787	.763	.739
8.74	6.67	.737	.716	.691	.667	.641
0	0	.644	.625	.601	.578	.556

Izbekov and Plotnikov, 1911

%	κ	%	κ
99.5°			
1.09	0.0006	60.19	116.46
3.60	.0530	65.29	132.12
5.68	.8540	71.70	146.10
7.31	2.6740	76.17	165.82
9.35	7.1480	81.06	178.08
22.31	36.0100	84.00	176.74
29.91	53.4800	92.86	117.41
40.00	73.0300	97.58	47.24
49.52	87.8200	99.02	20.86

Aluminum bromide ($AlBr_3$) + Bismuth bromide ($BiBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.	
0	97.1	51.9	153.1	
4.4	96.8	57.7	147.8	
10.6	95.5	65.6	137.3	
10.6	98.5	71.5	156.9	
23.6	119.2	79.7	180.1	
32.5	135.5	88.0	202.7	
43.5	150.2	100.0	220.4	(1+1)
46.6	152.0			

Izbekov, 1925

mol%	f.t.	mol%	f.t.
0	98.4	40.2	142.2
6.6	95.7	49.0	148.7
11.4	94.6	52.4	147.1
16.3	93.5	59.8	134.9
22.1	107.7	68.2	133.8
28.9	121.8	77.9	165.0
31.4	126.2	82.3	182.4
36.0	135.8	100	218.8
	(1+1)		

Aluminum bromide($AlBr_3$) + Stannic bromide ($SnBr_4$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.		mol%	f.t.
0	97.1		56.7	45.8
1.9	96.0	I	71.2	29.7
2.5	94.3		74.4	25.5
13.4	85.4		76.0	22.3
22.7	76.4		82.5	23.4
35.4	65.7	II	93.3	27.6
49.2	53.7		100	31.0

Pushin and Makucz, 1938

mol%	f.t.	E	mol%	f.t.	E
0	95	-	61.4	34.5	20
11.4	83	-	74.1	23	20
25.5	69	12	86.5	25	18
41.1	52.4	16	100	30	-
51.6	43.5	19			

E : 26.0 mol% - 20°

Antimonium tribromide ($SbBr_3$) + Bismuth tribromide ($BiBr_3$)

Pushin and Makucz, 1938

mol%	f.t.	E	mol%	f.t.	E
100	218	-	44.5	169	-
88.5	204	-	40	165	94
80.5	197	-	46.5	163	-
69	188	-	22.5	145	-
60	180	94	11	120	-
54.5	177	-	0	94	
52	174	94			

Antimonium tribromide ($SbBr_3$) + Stannic bromide ($SnBr_4$)

Pushin and Makucz, 1938

mol%	f.t.	m.t.	mol%	f.t.	m.t
0	94	-	59.7	70	-
9.9	90	24.5	69.6	64.5	26.2
22.3	86	25	79.5	55.5	27
29.9	83	25	89.5	36	27
40.9	78.5	25.3	94.2	24.2	27
50.8	74	25.7	100	30	-

E : 6.0% - 27.0°

Bismuth tribromide ($BiBr_3$) + Stannic bromide ($SnBr_4$)

Pushin and Makucz, 1938

mol%	f.t.	E	mol%	f.t.	E
0	218	-	37.5	211	27
7.5	212	-	54.5	211.5	30
13.4	-	28.5	88.5	212	28.5
16	211.5	-	95	211	-
29.3	211.8	29	100	30	-

Stannic bromide ($SnBr_4$) + Aluminum antimonium tribromide ($SbBr_3.AlBr_3$)

Gorenbein, 1951

%	d 85°	90°	95°	100°
0	-	3.168	-	3.139
26.7	3.199	3.180	3.165	3.151
34.0	3.209	3.194	3.179	3.164
43.9	3.232	3.217	3.203	3.188
53.9	3.254	3.238	3.225	3.211
63.9	3.279	3.263	3.248	3.233
74.8	3.296	3.282	3.269	3.255
88.4	3.324	3.311	3.299	3.286
90.9	3.327	3.316	3.306	3.295
95.0	3.334	3.322	3.310	3.299
100	3.349	3.338	3.326	3.315

%	η 85°	90°	95°	100°
0	-	1306	-	1223
26.7	1810	1710	1682	1540
34.0	2188	2016	1887	1765
43.9	2903	2626	2415	2236
53.9	4015	3613	3266	2996
63.9	5630	5034	4420	4000
74.8	8400	7400	6503	5724
88.4	13800	12000	10400	9200
90.9	15994	13455	11636	10080
95.0	19537	15609	13507	11532
100	28466	19561	16399	13957

%	κ 85°	90°	95°	100°
26.7	1.4	1.2	0.95	0.77
34.0	6.9	6.6	6.2	5.70
43.9	19.5	20.4	20.6	20.70
53.9	33.7	37.0	39.8	40.70
63.9	50.5	56.1	58.9	62.10
74.8	63.4	70.0	75.8	82.70
88.4	72.5	82.5	93.5	103.50
90.9	74.0	84.3	96.0	106.50
95.0	77.3	87.7	98.7	110.50
100	80.3	91.5	102.5	115.50

Lithium iodide (LiI) + Silver iodide (AgI)

Sandonnini and Scarpa, 1913

mol%	f.t.	tr.t.	mol%	f.t.	tr.t.
0.0	450	-	60.0	498	-
10.0	432	-	70.0	510	280
20.0	422	-	80.0	530	-
30.0	438	400	90.0	539	198
40.0	467	380	100.0	546	145
50.0	480	360			

Sodium iodide (NaI) + Potassium iodide (KI)

Greiner and Jellinek, 1933

mol% L	V	p_1	p_2
	1180°		
0	0	260	0
25	24.2	189.1	60.3
50	46.3	132.9	115.1
75	72.7	66.8	177.6
100	100	0	231

Kurnakov and Zhemchuzhni, 1907

mol%	f.t.	mol%	f.t.
100.00	693	44.14	586
87.84	669	34.10	589
75.06	639	23.70	607
64.36	614	10.14	637
54.63	598	0.00	660
50.00	591		

Zhemchuzhni and Kurnakov, 1910

mol%	Q melt. of mixed crystals
25	740
50	628
75	457
88	104
after 2 1/2 months	
50	400

Sodium iodide (NaI) + Silver iodide (AgI)

Sandonnini and Scarpa, 1913

mol%	f.t.	E	min.	tr.t.	min.
0.00	662	-	-	-	-
5.0	650	-	-	-	-
10.0	640	380	-	10	-
20.0	595	393	146	20	90
30.0	560	390	140	60	100
40.0	510	390	146	70	100
50.0	475	394	146	110	140
60.0		394	146	120	120
70.0	402	392	146	90	130
80.0	430	390	146	50	150
90.0	402	394	146	10	160
95.0	530	393	146		200
100.0	546	-	145	-	250

Zakharchenko and Bergman, 1955

%	f.t.	%	f.t.
0	632	60	452
10	630	65	432
20	596	70	412
25	580	73	399
30	562	75	390
35	544	80	424
40	527	85	460
45	509	90	492
50	490	95	524
55	472	100	554

Kolotii, 1956.

mol%	e 690°	e 700°	e 710°
0.23	-	-	0.4239
0.83	-	0.3884	0.3915
1.64	-	0.3187	0.3204
4.76	-	-	0.2244
9.09	-	0.1232	0.1395
16.69	-	-	0.1080
33.23	0.0990	0.1008	0.1032
45.60	0.0470	0.0455	0.0451
62.63	0.0392	0.0385	0.0380

mol%	e 720°	e 730°
0.23	0.4255	0.4250
0.83	0.4033	0.4014
1.64	0.3226	0.3246
4.76	0.2247	0.2254
9.09	0.1600	0.1741
16.69	0.1081	0.1111
33.23	0.1031	-
45.60	0.0456	0.0445
62.63	0.0367	-

Sodium iodide (NaI) + Magnesium iodide (MgI_2)

Klemm, Beyersdorfer and Oryschkewitsch, 1949 (fig.)

mol%	f.t.	mol%	f.t.
0	650	60	520
10	625	70	560
20	545	80	590
30	470	90	625
39	425 E	100	650
50	475		

Sodium iodide (NaI) + Cadmium iodide (CdI_2)

Brand, 1912

mol%	f.t.	E	min.
100	385	-	-
90	374	285	80
80	356	285	140
70	334	287	240
60	305	"	360
50	308	"	400
40	410	"	320
30	504	"	230
20	569	285	170
10	619	"	110
0	653	-	-

Ilyasov and Bergman,

mol%	f.t.	mol%	f.t.
100	388	52.2	287 E
90.5	375	51.0	306
81.9	363	48.2	328
74.0	351	45.5	354
66.7	332	43.9	371 tr.t
60.0	313	42.9	386
56.9	306	38.0	435
53.9	292	33.3	474
(2+1)		31.2	492

Sodium iodide (NaI) + Mercuric iodide (HgI_2)

Belayev and Mironov, 1952

mol%	f.t.	E	tr.t.	
100.0	257.5	127	-	
96.1	255.0	-	-	
95.2	254.0	125	-	
92.5	253.0	-	-	
91.7	252.0	120	217	
89.2	249.0	-	-	
88.0	248.0	121	216	
86.1	246.0	-	-	
85.0	242.0	-	-	
84.0	241.0	-	-	
81.8	239.0	120	216	
79.8	237.0	120	216	
78.0	232.0	-	-	
76.8	238.0	-	-	
76.2	226	-	-	
75.0	220.5	120	216	
74	-	-	216 E	
73.4	228.0	-	-	
71.8	238.5	-	-	
70.3	253.0	120	216	
69.8	269.5	-	-	(1+2)
66.4	294.0	216	256	
64.3	317.0	-	-	
62.3	338.0	-	-	
60.4	349.0	-	-	
58.6	367.0	-	-	
57.0	378.0	-	-	
55.3	390.0	216	256	
53.8	392	-	-	

Sodium iodide (NaI) + Lead iodide (PbI_2)

Ilyasov and Bostandzhiyan, 1956.

mol%	f.t.	mol%	f.t.
100	408	56.9	457
90.5	402	53.9	473
81.9	389	51.0	490
74	382	48.2	508
71.8	378 E	42.9	540
70.2	386	40.4	553
66.7	402	38	564
63.3	419	33.4	582
60	437	29.1	590

Potassium iodide (KI) + Magnesium iodide (MgI_2)

Klemm, Beyersdorfer and Oryschkewitsch, 1949 (Fig.)

mol%	f.t.	mol%	f.t.
0	685	60	510
10	620	70	570
20	535	80	610
30	420	90	635
39	255	100	650
50	420		

Potassium iodide (KI) + Lead iodide (PbI_2)

van Klooster and Stearns, 1933

mol%	f.t.	E	mol%	f.t.	E
100	412	-	47.5	-	346
95	404	321	45	378	"
90	395	321	40	422	"
85	383	"	35	475	"
80	367	"	30	504	"
75	349	"	25	548	"
70	324	"	20	585	•
69	-	"	15	609	347
65	330	"	10	641	-
60	337	"	5	660	-
55	345	"	0	686	-
50	349	-			

(1+1)

Ilyasov and Bergman, 1956.

mol%	f.t.	mol%	f.t.
100.0	408	60.0	349
90.5	391	56.9	353
81.9	376	53.9	356
78.8	364	51.0	358
74.0	347	48.2	362
70.2	327	42.9	405
66.7	"	38.0	455
63.3	342	33.3	503

(1+1)

Potassium iodide (KI) + Cadmium iodide (CdI_2)

Brand, 1912

mol%	f.t.	E	min.
100	385	-	-
90	370	185	100
80	351	186	220
70	323	186	370
60	286	187	450
55	247	185	470
53	197	185	500
51	195	185	450
50	203	186	420
47	231	188	360
45	242	188	330
40	302	188	210
35	383	188	160
33.7	395	188	120
30	440	178	70
20	560	176	60
10	634	175	-
0	678	-	-

(2+1) tr.t.=215°

Potassium iodide (KI) + Aluminum iodide (AlI_3)

Nijnik, 1937

%	f.t.	E	%	f.t.	E
100	186.5	-	100	188.3	-
98.68	178.7	-	92.07	167.8	97.8
95.21	174.7	98.4	89.58	160.8	101.0
89.84	158.3	97.4	88.36	149.0	97.8
84.50	111.6	97.5	86.62	130.0	97.8
78.69	159.6	97.6	85.64	-	98.0
76.32	176.3	97.3	83.64	110.8	98.0
73.00	196.6	97.4	82.63	112.6	98.4
			81.08	133.6	97.6
			68.56	-	206.9
			63.55	-	207.0
			60.85	-	206.9
			59.22	-	207.0

Izbekov and Nijnik, 1937

%	κ	%	κ
99.66	2.205	88.86	430.3
99.35	5.516	87.78	476.7
98.72	20.240	86.90	516.1
98.47	26.770	84.00	654.6
97.85	45.920	80.68	794.5
96.68	84.210	77.71	984.7
93.41	222.300	75.66	1116.0
89.72	390.700	73.85	1211.0

%	d
200°	
100	3.256
95.64	.213
89.68	.267
81.55	.337

Tubandt, Rheinhold and Jost, 1927

mol %	κ	mol %	κ
	410°		
0	80	70	208
5	78	90	226
10	85	95	229
30	119	100	236
50	155		

Barth and Lunde, 1926

Lattice constant

Lunde and Rosbaud, 1930

Crystal structure

Cuprous iodide (CuI) + Silver iodide (AgI)

Quercigh, 1914

%	f.t.	m.t.	tr.t. I	tr.t. II	min.
0	602	-	-	402	60
12.04	559	548	370	354	-
23.55	533	520	350	334	-
34.56	516	500	335	317	-
45.10	500	492	319	-	-
55.20	490	485	298	-	-
64.89	495	490	278	-	-
74.20	506	495	270	-	-
83.13	521	510	245	-	-
91.73	535	524	205	-	-
95.90	544	536	175	147	70
100	557	-	147	147	130

Cuprous iodide (CuI) + Cadmium iodide (CdI_2)

Herrmann, 1911

%	f.t.	m.t.	tr.t.	min.
0	606.5	-	414	8
20	564	592.2	378	6
50	474.5	440	290	3.9
70	413	395	272	2.5
90	350	-	264	0.5
92.5	361	250	260	-
95	371	350	-	-
100	392.4	-	-	-

Cuprous iodide (CuI) + Lead iodide (PbI_2)

Greiner and Jellinek, 1933

mol% L	mol% V	p_1	p_2
	800°		
100	100	0	355
90	-	-	305.4
70	96.4	8.0	209.7
50	91.7	14.4	153.7
30	80.5	20.7	88.4
10	42.3	33.8	24.5
0	0	42.6	0

Cuprous iodide (CuI) + Aluminum iodide (AlI_3)

Izbekov and Nijnik, 1937

%	κ	%	κ
99.15	1.270	94.00	27.07
97.83	6.657	92.62	47.79
96.69	20.040	85.89	55.67

Silver iodide (AgI) + Lead iodide (PbI_2)

Germann and Metz, 1931

mol%	f.t.	m.t.	
95	405	-	mixed cryst.
90	398	-	"
85	389	342	
80	383	344	
75	376	346	
70	368	348	
65	360	350	
60	353	350	mixed cryst.
57.50	350	-	"
55	351	-	"
50	356	-	"
45	362	-	"
40	368	-	"
35	375	-	"
30	382	-	"
25	390	-	"
20	399	-	"
18.50	402	-	(5+1)
16.66	408	402	"
15.75	420	402	
15	425	402	
10	467	402	
5	507	402	

Silver iodide (AgI) + Mercuric iodide (HgI_2)

Roozeboom, 1900 - 1901

mol%	f.t.	tr.t.
0	526	-
0-82	mixed cryst.	-
33	(2+1)	157
-	-	135
86	242 E	-
-	-	118
96-100	mixed cryst.	-
100	257	127

Steger, 1903

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	257	257	50	357	318
97	254	248	40	387	346
93	251	242	30	320	366
90	248	242	26	433	378
87	246	242	20	454	414
82	257.5	242	10	480	470
80	264	246	5	507	492
70	295.5	265	0	526	526
60	327	290			

mol%	I	tr.t.	II
100	127	- (1+2)	-
95	129-132	-	-
90	128.5-131.5	-	-
85	132-133	118-119	44-46
80	129-133	116-119	-
75	132-133	117-119	43.5-46
70	132-133	117-119	-
65	-	117-119	44-46
60	-	117-119	-
55	-	117-119	43-46
50	-	-	-
45	139.5-144	-	42-46
40	147-150	-	-
37	150-152.5	-	-
34	155-157.5	-	-
30	154-157	-	46-50
25	145-147	-	-
20	140-142	-	44-48
15	135-137	-	-
10	135	-	-
7.5	134-135	-	44-46
5	134-140	-	-
2.5	135-144	-	45-47
0	147	-	-

Bergman and Gönke, 1926

mol%	f.t.	E	I	tr.t. (2+1)	II
100	257	-	127	-	-
90	247	245	130	119	-
82	267	245	131	118	-
74	288	245	130	117	44
67	308	245	-	118	45
60	327	238	-	118	43
50	354	218	-	132	44
41	390	198	-	145	45
33	410	180	-	152	48
25	450	155	-	144	52
14	490	-	135	135	46
0	556	-	148	-	45

Silver iodide (AgI) + Lead iodide (PbI_2)

Tubandt and Eggert, 1920

mol%	f.t.	E	tr.t. I	tr.t. II
0	552	-	144.6	-
5	507	395	140	115
10	465	395	139.5	115
15	429	395	137	114
20	393	(393)	-	112
25	381	-	-	110
30	374	-	-	112
40	359	-	-	110
50	350	-	-	110
60	344	-	-	112
70	358	-	-	110
80	373	-	-	(110)
90	388	-	-	-
100	402	-	-	-

(4+1)

Silver iodide (AgI) + Aluminum iodide (AlI_3)

Izbekov and Nijnik, 1937

%	ϰ	%	ϰ
98.76	3.572	85.99	190.6
96.38	22.860	82.19	274.9
93.40	65.130	81.32	282.5
89.28	45.400	81.13	287.0

Cadmium iodide (CdI_2) + Mercuric iodide (HgI_2)

Sandonnini, 1912

mol%	f.t.	tr.t.	mol%	f.t.	tr.t.
0	380	-	60	309	-
10	366	-	70	300	105
20	345	-	80	290	118
30	334	-	90	270	124
40	325	-	100	253	128
50	314	-			

Losana, 1926

%	f.t.	m.t.	tr.t.
99.0	254	252	130
98.0	256	253	129
97.0	257	254	128.3
95.0	259	255	128
90.0	263.5	258	127
86.5	267.5	263	124
83.5	269	266	120
80.0	272	269	116
77.0	276	272	110
75.0	278	275	108
72.5	283	579	105
70.0	285	281	102

Cadmium iodide (CdI_2) + Aluminum iodide (AlI_3)

Izbekov and Nijnik, 1937

%	ϰ	%	ϰ
97.77	0.336	85.07	61.41
96.94	0.715	84.80	67.86
94.03	2.048	80.71	81.19
93.66	9.880	79.66	86.54
91.52	16.460	75.76	94.35
90.48	27.360	74.83	99.94

Mercuric iodide (HgI_2) + Lead iodide (PbI_2)

Losana, 1926

%	f.t.	m.t.	E	min.	tr.t.
0	253	-	-	390	131
1.16	252	250	-	-	128
3.03	251	246	-	-	124
5.30	249.2	240	-	-	122
8.06	246.5	235.8	-	-	119
10.00	244	233	-	-	117
12.00	242	232	-	-	115
14.00	240	230	-	-	113
16.54	237	229	-	-	112
19.88	232	-	228	60	129
21.54	230	-	227	190	130
23.08	-	-	228	370	-
25.00	234	-	226	340	-
30.00	249.8	-	229	220	-
40.00	281	-	228	180	-
50.00	307	-	226	160	-
60.00	320	-	229	100	-
70.00	348	-	226	60	-
72.86	352	-	228	30	-
74.18	354	-	-	-	-
75.00	355	266	-	-	222
80.00	363	303	-	-	162
85.00	368	327	-	-	129
88.16	374	348	-	-	-
90.00	376	350	-	-	-
93.00	378	360	-	-	-
95.00	380	365	-	-	-
100.00	393	-	-	300	-

Mercuric iodide (HgI_2) + Aluminum Iodide (AlI_3)

Nijnik, 1937

%	f.t.	E	%	f.t.	E
100	187.9	-	100	188.1	-
96.43	177.9	-	58.13	125.5	119.8
92.97	174.5	-	52.79	146.6	120.7
89.72	171.4	-	51.07	147.9	-
84.80	168.4	-	49.42	155.8	-
76.73	161.2	-	47.01	139.4	135.0
69.32	142.8	120.0	45.04	135.0	-
65.95	127.6	120.7	42.95	133.9	-
62.65	121.0	120.3	41.31	133.8	-
59.52	124.7	120.0	39.63	133.4	-
56.96	135.8	120.0	37.21	133.8	-
50.25	151.6	119.8	34.96	132.7	-
			30.60	-	133.3
0	251.0	-	77.67	160.5	-
23.53	217.4	135.0	63.26	122.3	-
33.52	194.5	133.9	63.72	122.6	-
40.08	-	135.7	48.93	156.3	-
47.76	144.0	133.6	48.35	156.9	
12.32	239.0	-			
19.19	226.8	-			

Izbekov and Nijnik, 1937

%	κ	%	κ
99.21	0.0558	66.27	281.2
97.08	0.3133	62.80	335.1
94.09	2.298	59.69	398.8
91.96	6.592	49.67	634.6
85.84	40.040	48.52	660.6
82.36	72.990	43.63	845.2
75.77	148.500	41.33	1490.0
69.38	236.600	22.85	2990.0

%	d	%	d
	200°		
100	3.256	66.65	3.820
86.57	3.436	52.49	4.194

Aluminum iodide (AlI_3) + Antimonium triiodide (SbI_3)

Nijnik, 1937

wt %	mol %	f.t.	E
0	0	187.7	-
15.56	13.01	168.4	140.0
22.35	18.24	162.6	140.3
29.20	25.06	160.0	140.5
36.82	32.11	151.6	140.2
42.38	37.39	145.4	140.1
48.36	43.19	147.8	135.8
59.52	54.40	-	135.1
61.43	56.62	142.9	-
52.47	47.23	149.6	140.2
60.04	55.07	144.7	135.9
72.46	68.13	142.9	135.8
87.21	84.69	151.6	135.8
95.67	94.72	160.4	-
100	100	168.6	
19.52	-	167.6	141.6
29.24	-	159.3	140.3
32.10	-	156.8	140.6
35.54	-	152.4	140.2
36.46	-	148.8	140.6
36.76	-	150.4	140.4
37.94	-	142.2	140.7
39.14	-	142.0	140.4
40.95	-	143.1	140.1
43.14	-	143.8	-
52.91	-	148.8	140.6
53.38	-	147.4	140.3
54.52	-	147.4	140.2
55.74	-	147.3	-
58.84	-	145.8	135.0
60.16	-	144.4	134.9
61.36	-	144.6	135.0
64.09	-	133.6	135.0
66.33	-	139.9	135.0
74.38	-	144.6	135.0
76.49	-	147.0	134.0

Izbekov and Nijnik, 1937

%	κ	%	κ
1.89	0.295	31.91	102.7
4.56	9.987	47.02	183.7
13.53	22.12	52.57	204.7
14.39	24.77	60.83	242.7
17.69	40.94	64.45	248.3
22.93	66.04	76.78	263.1
26.52	82.87	94.23	75.24
29.04	93.40		

%	d
200°	
0	3.256
11.42	.286
27.58	.440
63.00	.785

Aluminum iodide (AlI_3) + Stannic iodide (SnI_4)

Nijnik, 1937

wt%	mol%	f.t.	E
0	0	188.0	-
11.70	7.98	173.0	-
20.93	14.77	163.3	114.0
31.90	23.45	152.5	113.8
38.17	29.13	144.6	113.6
45.67	35.48	133.8	113.8
48.73	38.35	129.0	113.8
53.45	42.91	125.8	114.4
58.50	48.09	122.9	113.7
63.78	53.52	118.9	113.8
67.32	57.41	116.3	112.9
67.26	57.32	117.8	113.5
74.48	65.62	119.5	114.6
82.50	75.47	123.3	113.4
91.49	86.74	135.9	116.2
100	100	142.6	-
75.31	66.65	119.8	113.7
69.97	60.39	113.6	113.6
80.10	72.48	129.3	113.8
67.26	57.32	117.8	113.5
67.32	57.41	116.3	112.9
69.97	60.39	113.6	113.6
74.48	65.62	119.5	114.6
75.31	66.65	119.8	113.7
80.10	72.48	129.3	113.8
82.50	75.47	123.3	113.4
91.49	86.74	135.9	116.2
100	100	142.6	-

Cuprous-mercuriiodide (Cu_2HgI_4) +
Silver-mercuriiodide (Ag_2HgI_4)

Suchow and Keck, 1953 (fig.)

mol%	t of colour change heating	cooling
100	51	51
80	37	42
60	34	35
57	34	34
48	34	39
40	39	48
20	53	62
0	68	68

61 - 46mol% , $C_1 + C_2$, below 42°

Sodamide ($NaNH_2$) + Potassamide (KNH_2)

Kraus and Cuy, 1923

mol%	f.t.	E	mol%	f.t.	E
0	206.4	-	38.0	100.3	97
6.4	189.5	-	43.2	110.5	92
14.0	155.0	92	54.9	118.5	97
18.2	145.1	92	64.1	120	-
25.1	120.5	92	72.0	174	120
32.9	97	97	82.4	231	120
34.3	97	97	100	329	-

(1+2)

Sodium cyanide(NaCN) + Potassium cyanide (KCN)

Truthe, 1912

%	f.t.	m.t.	tr.t.	min.
0	561.7	-	-	-
7.0	558	536	-	-
15.0	538	517	-	-
30.0	512	502	259	5.0
40.0	510	495	261	7.1
50	502	502	263	7.7
60.0	518	504	261	5.1
70.0	541	517	260	4.0
85.0	575	533	-	-
95.0	610	588	-	-
100	622	-	-	-

Sodium cyanide (NaCN) + Cuprous cyanide (CuCN)

Truthe, 1912

%	f.t.	m.t.	E	min.	tr.t.	min.
100	473	-	-	-	-	-
90	408	-	345	25.0	-	-
85	357	-	347	27.7	-	-
82.5	353	-	345	36.0	-	-
78.5	371	-	345	23.7	-	-
75	381	-	345	14.7	-	-
70	395	-	343	4.2	-	-
65	398	-	-	-	-	-
62	389	-	371	1.3	318	4.3
60	382	-	372	3.2	318	7.3
55	373	366	-	-	318	10.0
52	375	372	-	-	318	6.1
50	377	375	-	-	-	-
45	396	381	-	-	-	-
40	406	392	-	-	-	-
30	450	408	-	-	-	-
25	483	430	-	-	-	-
20	505	459	-	-	-	-
10	537	508	-	-	-	-
0	562.3	-	-	-	-	-

(1+2)

Sodium cyanide (NaCN) + Silver cyanide (AgCN)
Truthe, 1912

%	f.t.	m.t.	E	min.
0	562.3	-	-	-
7	561	540	-	-
15	553	500	-	-
23	536	-	422	10.1
30	521	-	419	14.0
40	490	-	421	22.6
50	443	-	424	39.8
53	-	-	424	31.3
60	446	-	423	17.0
67	-	-	-	-
73.2	471	-	-	-
100	520-350	-	422	-

Potassium cyanide (KCN) + Cuprous cyanide (CuCN)

Truthe, 1912

%	f.t.	E	min.	tr.t.	min.
100	473	-	-	-	-
95	447	280	-	228	1.4
85	317	280	-	230	4.7
82	-	278	28.6	229	5.1
80	283	279	23.4	230	4.8
75	312	280	10.2	230	0.9
72	327	-	-	-	-
70	324	-	-	-	-
67.5	324	-	-	-	-
65	325	-	-	-	-
60	327	-	-	-	-
55	315	272	15.4	228	4.3
50	288	277	40.0	227	7.8
48	288	275	33.6	228	7.7
45	335	277	24.1	227	7.2
40	386	275	10.5	230	5.3
35	399	277	4.4	224	2.3
32.5	400	-	-	-	-
30	424	398	47.7	-	-
20	532	398	25.0	-	-
10	586	387	19.0	-	-
0	622.5	-	-	-	-

(1+2)

Potassium cyanide (KCN) + Silver cyanide (AgCN)

Truthe, 1912

%	f.t.	E	min.	
0	622.5	-	-	
10.0	608	290	8.1	
20.0	589	290	13.3	
30.0	541	293	16.2	
40.0	462	289	24.7	
47.5	356	290	36.2	
50.0	333	291	45.1	
52.5	289	289	55.5	
54.0	298	292	50.4	
60.0	333	291	33.3	
65.0	360	291	10.8	
67.4	370	-	-	
70.0	368	292	1.2	
84.0	338	291	7.5	
90.0	313	289	15.0	(1+1)
92.0	307	290	20.8	
100.0	320	-	-	

Potassium cyanide (KCN) + Zinc cyanide (ZnC_2N_2)

Truthe, 1912

%	f.t.	E	min.
100	-	-	-
90	-	512	11.6
80	-	511	15.1
70	-	516	18.0
60	518	515	28.0
50	535	515	6.6
47.4	538	-	-
45	535	486	-
40	525	487	15.0
30	-	487	43.8
20	559	486	25.6
10	602	485	11.5
0	622.5	-	-

(1+2)

Potassium-zinc cyanide ($K_2ZnC_4N_4$) + Potassium-cadmium cyanide ($K_2CdC_4N_4$)
Carozzi, 1926

%	d	n_D	%	d	n_D
		25°			
100	1.824	1.413	39.98	1.744	1.4175
79.24	1.789	1.4141	18.87	1.682	1.4191
60.59	1.768	.416	0	.673	.4213

Potassium-zinc cyanide ($K_2ZnC_4N_4$) + Potassium-mercuric cyanide ($K_2HgC_4N_4$)

Carozzi, 1926

%	d	n_D	%	d	n_D
		25°			
100	2.438	1.413	23.92	1.801	1.447
46.79	1.962	.436	11.64	.738	.452
34.58	1.883	.443	0	.673	.458

Potassium-cadmium cyanide ($K_2CdC_4N_4$) + Potassium-mercuric cyanide ($K_2HgC_4N_4$)

Carozzi, 1926

%	d	n_D	%	d	n_D
		25°			
100	2.438	1.421	38.71	2.005	1.441
78.00	.261	.428	17.90	1.927	.451
60.60	.151	.435	0	1.824	.458

Sodium thiocyanide (NaCNS) + Potassium thiocyanide (KCNS)

Wrzesnewsky, 1911 and 1912

mol%	f.t.	E	tr.t. I	tr.t. II
100	179.0	-	143.0	-
95	167.5	120	142.5	102
90	160	121.5	142.4	103
85	152.5	122	141.9	102
80	138.5	124	-	102
75	133.5	123	-	102 4
70	123.5	123.5	-	102
66	161	122	-	103
60	190.5	125	-	102
55	205.5	125.5	-	-
50	229	122.5	-	-
40	250	123	-	-
34	263.8	122	-	-
30	272	121	-	-
20	292	121	-	-
10	306	120.5	-	-
5	315	-	-	-
0	323	-	-	-

mol%	flowing pressure
100	19.0
95	46
90	64
80	79
75	74
70	56
60	50
50	53
40	59.1
34	61
20	71
10	77.2
0	83.5

Rubidium thiocyanide (RbCNS) + Potassium thiocyanide (KCNS)

Wrzesnewsky, 1911 and 1912

mol%	f.t.	m.t.	tr.t.
100	195	-	-
95	191	190	-
90	188	185.5	-
80	183	178.8	-
70	179	175	126
60	176.4	171.3	136
55	175	169	135.5
50	174	166	136
45	172.1	165	136
40	170.8	165	135
30	171.4	165	134
20	172	167	133
15	172	168.5	134.6
5	177	175	139.5
0	179	-	143

mol%	flowing pressure
100	39
90	57.8
80	68
70	63.8
60	47.9
50	37.8
40	29.9
30	25.2
20	23
10	20.2
0	18.5

Cadmium-mercuric thiocyanate ($CdHgC_4N_4S_4$) + Cobalt-mercuric thiocyanate ($CoHgC_4N_4S_4$)

Straumanis and Stahl, 1943

Lattice constants.

mol%	d	a in Å	c in Å
	25°		
0	3.2731	11.4403	4.2043
1.87	-	11.4312	4.2038
3.67	3.2676	(11.426)	(4.208)
48.81	3.1532	11.241	4.296
65.60	3.1111	(11.206)	(4.310)
95.02	-	11.096	4.3643
100.00	3.0247	11.0868	4.3652

Ammonium chlorostannate ($N_2H_8SnCl_6$) + Ammonium chloroplumbate ($N_2H_8PbCl_6$)

Carozzi, 1926

%	d
18°	
100	2.925
80.20	2.842
54.40	2.694
17.90	2.468
0	2.390

%	n_D
18°	
80.2	1.74
54.4	1.715
17.9	1.69
0	1.675

Ammonium chlorostannate ($N_2H_8SnCl_6$) + Ammonium chloroplatinate ($N_2H_8PtCl_6$)

Carozzi, 1926

%	d
	18°
100	3.009
99.70	3.006
84.80	2.881
78.00	2.810
25.50	2.524
0	2.390

Diphenylmercury ($HgC_{12}H_{10}$) + Triphenyl bismuth ($BiC_{18}H_{15}$)

Cambi, 1912

mol%	f.t.	E	min.
100	76	-	-
85.2	72.8	63.5	10
72.4	66.6	63.6	80
66.2	64.5	64.5	180
61.3	72.4	63.6	160
50.8	83	63.4	120
39.6	92.4	62.8	100
30.4	101	62.6	60
21.9	106.4	62.6	20
14.1	111.5	-	-
7.7	116.2	-	-
0	121.8	-	-

Diphenylmercury ($HgC_{12}H_{10}$) + Triphenylstibine ($SbC_{18}H_{15}$)

Cambi, 1912

mol%	f.t.	E	min.
100	49.8	-	-
91.8	45.6	40.6	40
80.2	41.8	41.5	200
70.1	61	40.4	140
60.2	68.5	41	140
50.2	86.2	39.8	60
40.3	95.3	40	30
30.2	102.5	39.6	10
20.1	110	-	-
10.1	116.7	-	-
0	121.8	-	-

Diphenylmercury ($HgC_{12}H_{10}$) + Tetraphenyltin ($SnC_{24}H_{20}$)

Cambi, 1912

mol%	f.t.	E	min.
0	121.8	-	-
1.6	117.8	117.8	240
8.4	134	112.6	230
17.2	156.8	114.6	180
35.6	180.5	114	170
55.4	199	112.6	60
65	209.5	110	40
76.8	214.2	108	-
82.4	216	-	-
88.2	218.8	-	-
100	223	-	-

Triphenylstibine ($SbC_{18}H_{15}$) + Tetraphenyltin ($SnC_{24}H_{20}$)

Cambi, 1912

mol%	f.t.	E	min.
100	223	-	-
90	221	-	-
80	218.4	47.8	10
70	210.6	46.2	30
60	204	46.4	50
50	194	46.2	70
40	186	47.8	120
25	162	47.5	140
10	122.8	49.2	220
6	95.8	49.8	260
0	49.8	-	-

Tetraphenyltin ($SnC_{24}H_{20}$) + Tetraphenyllead ($PbC_{24}H_{20}$)

Pascal, 1912

%	f.t.	m.t.	%	f.t.	m.t.
0	225.7	225.7	62.50	227.1	225.8
16.66	226.0	225.6	80	227.3	226.3
35	226.3	225.6	90.91	227.6	226.6
50	226.8	225.7	100	227.7	227.7

Lithium oxide (Li_2O) + Germanium oxide (GeO_2)

Schwartz, 1929

%	f.t.	m.t.	E	
63.6	1298	-	-	
67.7	-	-	1115	
69.4	1165	-	1114	
71.4	1128	-	1115	
73.7	1191	-	1114	
77.2	1239	1239	-	
81.3	1201	-	898	
83.2	1159	1161	884	
85.4	1100	-	897	(1+1)
87.5	1036	1036	886	
89.8	940	-	920	
92.1	-	-	920	
94.6	961	-	917	
100	1126			

Budnikov and Tresvyatski , 1954

mol%	E	f.t.	tr.t.
100	-	1115	-
95	920	1030	1025
90	920	1080	1035
85	925	1050	1035
80	925	1000	-
78	920	968	-
75	920 E	-	-
70	920	-	-
65	920	1050	-
60	920	1135	-
55	930	1200	-
50	-	1237 (1+1)	-
45	1115	1215	-
40	1120 E	-	-
35	1115	1270	-
33.4	-	1295 (1+2)	-
30	1125	1285	-

Lithium oxide (Li_2O) + Molybdenum oxide (MoO_3)

Hoermann, 1929

mol%	f.t.	E	min.	tr.t.	min.	
82.7	705	-	-	-	-	
84.4	680	511	65	-	-	
86.1	652	517	130	-	-	
87.9	616	523	180	-	-	
88.7	595	525	185	-	-	
89.6	570	529	230	-	-	
90.4	541	599	245	-	-	
90.8	529	532	255	-	-	
91.3	533	530	210	-	-	
92.2	541	-	130	-	-	(1+2)
93.0	546	-	-	-	-	
93.9	551	-	-	54[illegible]	110	(1+3)
94.2	558	-	-	548	60	"
94.8	563	-	-	546	-	"
95.3	568	-	-	568	180	(1+4)
96	584	-	-	566	150	
96.5	615	-	-	562	125	
97.5	632	-	-	553	100	
98.2	682	-	-	548	55	
100	732	-	-	-	-	
	795	-	-			

Lithium oxide (Li_2O) + Tungsten oxide (WO_3)

van Liempt, 1925

mol%	f.t.	E	tr.t.	
50	742	-	-	
52.5	722	694	-	
55.5	696	696	-	
59	710	700	-	(1+1)
62	735	698	-	
66.5	745	-	687	
69	740	-	661	
71	753	-	-	
77-100	two phases $L_1 + L_2$			

Hoermann, 1929

mol%	f.t.	E	min.	tr.t.	min.	
88.5	742	-	-	-	-	
89.6	721	694	120	-	-	
90.8	695	655	245	-	-	
92	718	695	150	-	-	
93.1	735	698	70	-	-	(1+2)
94.2	745	-	-	-	-	
94.5	744	740	132	-	-	
94.3	740	740	276	-	-	
95.4	784	740	220	-	-	
96.2	870	738	165	800	20	(1+4)
96.5	-	738	155	800	30	
97.1	-	738	95	799	50	
97.7	-	740	60	800	25	
100	1473	-	-	-	-	

Lithium oxide (Li_2O) + Vanadium pentoxide (V_2O_5)

Canneri, 1928

mol %	f.t.	E	min.
100	675	-	-
90.9	660	-	-
78.1	630	-	-
71.3	610	-	-
62.4	580	-	-
58.8	564	564	120
55.5	580°	564	60
52.6	600	564	30
50.0	618	-	-
47.6	615	-	-
44.4	608	-	-
40.0	590	515	60
38.4	585	575	60
37.0	582	575	90
35.7	575	575	90
34.4	850	575	60
33.3	950	-	-
31.2	855	-	-
	(1+1)	(2+1)	

Sodium oxide (Na_2O) + Barium sesquioxide (Ba_2O_3)

Bresker and Evstropiev, 1952

%	d	n_D	%	d	n_D
		20°			
100	1.8343	1.4625	83.47	2.1567	1.5000
98.20	.8901	.4715	82.70	.1713	.5009
97.21	.9099	.4744	81.55	.1867	.5021
96.34	.9332	.4772	80.87	.1999	.5033
95.43	.9535	.4806	79.88	.2146	.5045
94.34	.9744	.4828	78.93	.2315	.5055
93.52	.9936	.4857	77.77	.2493	.5072
92.62	2.0113	.4876	76.93	.2692	.5095
91.61	.0323	.4903	76.00	.2835	.5111
90.67	.0466	.4915	74.88	.3018	.5127
89.78	.0623	.4928	74.09	.3141	.5140
88.12	.0752	.4940	73.34	.3277	.5152
87.22	.0912	.4953	72.50	.3396	.5160
86.22	.1174	.4972	71.71	.3488	.5166
85.17	.1299	.4978	70.40	.3591	.5172
84.45	.1425	.4988	68.37	.3755	.5174

Sodium oxide (Na_2O) + Vanadium pentoxide (V_2O_5)

Canneri, 1928

mol%	f.t.	E	min.
100	675	-	-
98.1	650	-	-
73.0	635	-	-
69.0	600	-	-
66.2	595	-	-
60.2	575	-	-
55.6	565	565	120
51.3	580	565	60
50.0	618	-	-
47.6	630	-	-
42.2	620	-	-
41.4	595	570	90
41.2	580	574	120
40.4	574	574	120
36.2	578	574	90
33.4	610	574	60
32.3	632	-	-
31.7	623	620	90
30.8	620	620	120
28.6	620	620	30
25.0	620	620	60
	-	-	-
(1+1)	(2+1)	(3+1)	

Sodium oxide (Na_2O) + Germanium oxide (GeO_2)

Canneri, 1928

mol%	f.t.	E	min.
100	675	-	-
90.9	660	-	-
78.1	630	-	-
71.3	610	-	-
62.4	580	-	-
58.8	564	564	120
55.5	580	"	60
52.6	600	"	30
50.0	618	-	-
47.6	615	-	-
44.4	608	-	-
40.0	590	565	60
38.4	585	575	60
37.0	582	"	90
35.7	575	"	90
34.4	850	"	60
33.3	950	-	-
31.2	855	-	-

(1+1) , (2+1)

Schwartz, 1929

%	f.t.	E	
		1	2
60.2	1055	860	-
62.7	1078	-	-
65.1	1054	723	743
67.8	1012	725	745
70.7	934	735	745
72.2	766	739	745
73.8	804	743	743
75.4	-	742	-
76.2	-	743	-
77.1	-	-	-
78.0	751	-	-
78.9	774	-	-
80.0	820	769	-
(1+1)	(1+2)	(1+3)	(1+4)

Sodium oxide (Na_2O) + Titanium oxide (TiO_2)

Junker, 1936

mol%	f.t.	
100	1800	
90	1580	
80	1320	
78	1300	(1+3)
73	1120	
70	1090	
60	970	

Lepinskikh, Esin and Shavrin, 1956

mol %				
	1200°	1190°	1180°	1170°
23.5	25400	-	22400	-
29.0	21000	19000	17100	-
34.0	17600	-	15400	-
42.0	15500	-	11500	-
49.0	11600	-	9400	-
52.0	10300	8600	7500	-
66.2	7400	6000	-	5100
	1160°	1150°	1140°	1130°
23.5	-	18200	-	15200
29.0	13600	-	11400	-
34.0	11000	-	10100	-
42.0	-	9200	-	8400
49.0	8000	7400	6700	-
52.0	6400	5700	-	-
66.2	-	4100	-	3500
	1120°	1110°	1100°	1090°
23.5	-	13000	11500	-
29.0	-	-	9400	-
34.0	9200	-	-	-
42.0	-	8000	-	-
49.0	-	-	5600	-
52.0	4900	-	4500	-
66.2	-	-	3200	-
	1080°	1070°	1050°	1040°
23.5	10600	-	-	9400
29.0	-	8800	8500	-
34.0	8000	-	7100	-
42.0	7100	-	-	5500
49.0	-	5100	-	-
52.0	4300	-	3700	-
66.2	3000	-	-	-
	1020°	1010°	1000°	980°
23.5	-	9300	8600	8000
29.0	7500	-	7000	-
34.0	6100	-	5600	-
42.0	-	5000	-	-
49.0	4400	-	3800	-
52.0	3400	-	3200	3000
66.2	-	-	-	-
	970°	960°	950°	940°
23.5	-	7600	-	-
29.0	6200	-	5800	-
34.0	5100	-	5000	-
42.0	3900	-	-	3600
49.0	3100	-	-	-
52.0	-	-	-	-
66.2				
	920°	910°	890°	
23.5	-	6500	6400	
29.0	5400	-	-	
34.0	4500	-	-	
42.0	3400	-	-	

Sodium oxide (Na_2O) + Molybdenum oxide (MoO_3)

Groschuff, 1908

%	f.t.	E	min.	
100	791	-	-	
97.0	732	487	175	
94	626	493	300	
92.5	561	491	360	
91	-	494	420	
90.3	-	499	440	
89.5	506	499	400	(1+2)
88	531	495	295	
85	582	489	130	
82.3	612	550	-	
82	607	550	-	
79	592	552	200	(1+4)
77.5	575	552	290	
76	554	549	380	
74.5	586	549	295	
73	623		190	
71.5	655	546	125	
70	686	-	-	

%	tr.t. 1	min.	2	min.	3	min.
73	614	10	567	10	412	15
71.5	620	50	570	15	410	22
70	609	85	575	20	408	30

Hoermann, 1929

mol%	f.t.	E	min.	
70.0	687	-	-	
73.0	643	552	106	(1+2)
76.0	593	555	205	
77.9	556	557	245	(1+3)
79	570	556	212	
80.5	587	556	132	(1+4)
82	600	553	85	

mol%	f.t.	t (1+3)	min.	t (1+4)	min.
85	612	507	105	-	-
88	596	511	175	-	-
89.5	576	515	225	-	-
90.7	562	519	140	-	-
91	546	524	95	512	36
91.6	541	528	-	512	80
92.5	530	-	-	515	195
93.1	522	-	-	514	-
94	514	-	-	-	-
95.5	563	-	-	-	-
97	632	-	-	-	-
100	700	-	-	-	-
	790	-	-	-	-

Zelikman and Gorovits, 1954

mol%	f.t.	mol%	f.t.
50	690	71	600
55	600	77	525
58	560	82	510
62	610	91	720
67	620	100	800

Morris, Cook and al., 1955

t	d	t	d
50 mol %		77 mol %	
825	2.68	809	3.05
53 mol %		776	3.09
		652	3.24
828	2.73	82 mol %	
58 mol %		746	3.14
828	2.79	704	3.18
794	2.79	675	3.21
751	2.79	665	3.22
698	2.79	648	3.24
60 mol %		89 mol %	
790	2.84	844	3.06
740	2.88	797	3.10
690	2.94	730	3.16
66 mol %		91 mol %	
816	2.90	881	3.01
71 mol %		834	3.07
		776	3.10
757	3.05	100 mol %	
687	3.09		
661	3.13	909	2.98
		832	3.08

t	κ(in mhos)	t	κ(in mhos) t
50 mol %			
964	1.940	804	1.432
885	1.699	751	1.260
856	1.605		
53 mol %			
953	1.782	800	1.348
905	1.622	752	1.192
851	1.486		
55.5 mol %			
827	1.508	751	1.266
760	1.290		
58.4 mol %			
828	1.560	698	1.050
751	1.258	649	1.875
60.2 mol %			
806	1.425	692	1.045
751	1.250	656	1.[illegible]44
714	1.111		
66.6 mol %			
793	1.185	700	1.898
762	1.088	676	1.825
760	1.085	654	1.768
709	1.954		
71 mol %			
827	1.263	674	0.787
774	1.024	659	0.762
686	0.830	640	0.703
77 mol %			
825	1.185	712	1.837
805	1.128	694	1.793
758	1.995	662	1.690
730	1.916	641	1.630
81.9 mol %			
802	1.086	756	1.930
796	1.071	693	1.730
792	1.054	640	1.543
790	1.039		
89.2 mol %			
817	0.767	805	0.773
815	0.767	714	0.649
809	0.765	708	0.637
90.9 mol %			
890	1.140	812	1.821
819	1.955	785	1.875
100 mol %			
914	1.092	797	1.844
833	1.915	835	1.920

Sodium oxide (Na_2O) + Tungsten oxide (WO_3)

Van Liempt, 1925

mol%	f.t.	E	tr.t.
50	705	-	588 - 577
51.2	685	-	556
52.6	671	-	547
54.06	653	612	515
55.5	628	620	493
57.1	645	628	-
58.8	678	630	483
60.6	697	633	-
62.4	717	632	-
64.5	727	623	-
66.6	738	629	547
71.4	731	-	-
100	-	-	-

Hoermann, 1928

%	f.t.	E	min.	tr.t.		min.
79.0	700	-	-	590-583	(1+1)	-
81.1	-	-	-	547	"	-
82.15	652	628	175	515	"	-
83.2	-	-	-	493	"	-
84.9	665	629	220	-	(1+1)	-
85.3	677	629	180	475	"	-
86.3	696	626	125	440	"	-
89.5	738	-	-	-		-
90.5	728	728	245	-		-
91.6	755	730	175	-		-
92.0	768	728	160	-		-
93.0	824	727	125	784	(1+2)	56
93.7	-	727	95	780	"	90
94.7	-	727	72	777	"	180
95.1	-	-	-	776	"	186
95.8	-	-	-	773	"	180
96.4	-	-	-	773	"	70
96.5	-	-	-	722	"	60

Potassium oxide (K_2O) + Molybdenum oxide (MoO_3)

Hoermann, 1929

%	.t.	E	min.	
60.0	921	-	-	
64.0	883	462	70	
68	827	465	125	
72	746	481	165	
76	643	483	205	
78	583	485	240	
80	516	492	264	
80.3	492	492	295	
82	520	490	210	(1+3)
86.4	552	479	120	
86.8	571	-	-	
88	570	537	40	
90	563	537	100	
92	549	536	190	
92.8	537	537	240	
94.0	606	534	195	
96	696	532	140	
100	795	-	-	

Amadori, 1913

mol %	f.t.	E
100	786	-
97.0	755	522
90.9	691	523
88.4	637	523
85.4	574	526
83.3	-	526
82.6	540	525
81.9	548	522
80.0	560	-
78.7	552	544
77.5	-	544
76.9	552	544
75.7	555	-
75.2	558	-
72.4	542	480
70.4	518	480
68.0	-	480
66.6	484	-
65.3	-	476
63.2	608	476
58.4	750	473
54.9	840	471
53.4	875	468
51.51	907	466
50	926	-
(1+4)	(1+2)	

Spitsin and Kulechov, 1951

mol%	f.t.		mol%	f.t.	
50	926	(1+1)	73	557	
51	901		74	567	
53	873		75.5	571	(1+3)
54	840		76	567	
56	803		77	565	
57	759		77.5	560	
59	702		78	555	
61	639		78.5	553	E
61.7	594		79	556	
62	577		80	559	(1+4)
62.5	560		80.5	556	
63	545		81	554	
63.5	524		82	547	E
64.5	467	E	82.5	564	
65	474		83	579	
65.3	480		84	609	
65.7	486		85	625	
66	488		85.5	640	
66.5	489	(1+2)	86	649	
67	487		87	669	
67.5	485		88	697	
68	481	E	91	724	
68.5	489		93	740	
69	498		95	764	
70.5	521		100	795	
71	541				

Potassium oxide (K_2O) + Tungsten oxide (WO_3)

Van Liempt, 1925

mol%	f.t.	E	mol%	f.t.	E
50	933	-	60.6	799	595
51.2	917	555	62.4	799	600
52.9	890	579	64.5	799	602
54	852	580	66.6	637	584
55.5	831	588	68.9	694	590
57.1	799	589	71.3	748	581
58.8	-	584	100	-	-

Hoermann, 1929

%	f.t.	E	min.	tr.t.	min.	
71.0	921	-	-	-	-	
79.7	749	593	210	-	-	
82.6	664	598	275	-	-	
84	614	600	315	-	-	
85.5	637	600	225	-	-	
87.5	697	548	120	660	50	(1+3)
88.4	740	543	90	660	115	
90.4	841	-	-	660	210	
91.3	877	-	-	657	160	
92.7	953	-	-	930	120	(1+4)
94.2	1020	-	-	915	80	
97.4	-	-	-	910	60	

Potassium oxide (K_2O) + Barium sesquioxide (Ba_2O_3)

Bresker and Evstropiev, 1952

%	n_D	d	%	n_D	d
		20°			
100	1.4625	1.8343	77.00	1.4873	2.1278
97.12	.4723	.8995	74.78	.4877	.1428
94.52	.4791	.9497	72.68	.4888	.1674
91.83	.4842	.9886	70.13	.4917	.1992
89.32	.4867	2.0227	68.00	.4947	.2280
86.72	.4884	.0495	65.69	.4971	.2537
84.44	.4888	.0710	63.33	.4999	.2783
81.80	.4885	.0900	61.28	.5013	.2951
79.42	.4882	.1088	58.94	.5021	.3026

Potassium oxide (K_2O) + Vanadium pentoxide (V_2O_5)

Canneri, 1928

mol%	f.t.	E	min.
100	675	-	-
87.0	625	-	-
83.3	600	-	-
77	550	-	-
69.5	520	-	-
69.0	500	-	-
67.6	475	-	-
62.9	375	-	-
60.2	350	350	120
57.1	385	350	90
54.0	430	350	30
56.2	437	-	-
42.2	460	440	60
45.5	540	440	120
43.5	520	440	90
41.1	625	440	60
39.6	725	-	-
37.3	820	-	-
36.4	875	-	-
35.7	910	-	-
33.4	900	-	-
32.3	870	-	-
30.8	860	825	60
28.6	825	825	120
27.4	855	825	90
26.2	900	-	-
25.0	975	-	-
24.7			
(1+1)	(2+1)	(3+1)	

Illarianov, Ozerov and Kildisheva, 1956 (fig.)

mol%	f.t.	E	tr.t. I	tr.t. II	tr.t. III
	672	-	-	-	-
100	640	528	-	389	-
95	620	528	-	385	-
88.75	615	528	-	382	-
87.50	600	528	-	382	-
85	528	528	-	382	-
83	540	528	-	382	-
80	530	-	488	382	-
79	510	-	488	382	-
75	497	-	488	382	-
72.5	488	389	488	382	365
70	485	389	-	-	365
65	445	389	-	-	365
62.5	410	389	-	-	365
61.25	389	389	-	-	365
60	410	389	-	-	365
55	460	389	-	-	-
50	497	-	-	-	-
	(1+4)		(2+5)		

Potassium oxide (K_2O) + Germanium oxide (GeO_2)

Schwartz and Lewinsohn, 1930

%	f.t.	E
52.6	823	-
58.1	787	708
59.7	776	710
64.9	729	710
68.9	783	-
70.0	-	731
73.5	882	736
76.4	996	-
77.0	1017	-
78.0	1027	-
80.0	1027	-
81.6	1033	-
82.1	1018	965
84.6	977	966
90.0	glass	-
100.0	1126	-
(1+1)	(1+2) (1+3)	(1+4)

Rubidium oxide (Rb_2O) + Molybdenum oxide (MoO_3)

Spitsin and Kulechov, 1951

mol%	f.t.	mol%	f.t.
50	929 (1+1)	68	496
51	909	68.4	506
52.5	873	70.4	517
54	837	71.3	532
55	796	72.6	545
57	754	74	556
59	705	75.3	560
61	644	76.3	563 (1+3)
62	611	76.9	562
61.2	587	77.5	558
62.4	570	78.1	555
62.8	546	78.7	552
63.2	524	79.3	549 E_3
63.6	500	80	552
64.1	458 E_1	80.6	554 (1+4)
64.5	460	81.2	553
64.9	465	81.9	550
65.3	470	82.6	546 E_4
65.7	473	83.3	554
66.2	477	84	562
66.6	479 (1+2)	84.7	569
67.1	478 E_2	85.4	583
67.5	487	86.2	604
68	496	86.9	610
68.4	506	87.7	624
70.4	517	89.2	655
71.3	532	90.9	692
		92.5	722
		95.2	750
		100	795

Cesium oxide (Cs_2O) + Molybdenum oxide (MoO_3)

Spitsin and Kulechov, 1951

mol%	f.t.	mol%	f.t.
50	925 (1+1)	72	542
51	902	74	544
52.5	866	75	545 (1+3)
54	827	76	543
55.5	785	77	539
57	734	77.5	535
59	679	78	523 E
61	624	78.7	529
62	567	79	532
62.5	553	80	534 (1+4)
63	529	80.5	532
63.2	506	81	530 E
63.5	491	82	539
64	470	82.5	546
64.5	458 E	83	550
64.9	472	84	560
65.3	479	84.5	567
65.7	484	85.5	571
66	488	86	575
66.6	494	87	593
67	499	89	630
67.5	505	91	678
68	509	93.5	708
69	517	95	742
70.5	527	100	795
71	537		

Thallium oxide (Tl_2O) + Vanadium pentoxide(V_2O_5)

Canneri, 1928

mol%	f.t.	E	min.
100	675	-	-
96.2	650	-	-
91.0	635	-	-
82.0	575	-	-
69.2	525	-	-
66.7	475	-	-
62.5	400	-	-
59.2	240	340	120
56.8	360	340	90
52.6	375	340	30
50.0	390	-	-
47.6	375	-	-
44.4	390	330	-
43.5	375	"	60
41.4	350	"	120
40.0	330	"	90
38.0	320	"	60
36.4	335	"	-
31.4	400	310	-
	390	310	-
	420	310	-
28.6	450	-	60
	470	-	120
26.0	515	-	90
	550	-	-
25.0	555	-	-
24.2	550	-	-

(1+1) (2+1) (3+1)

Cuprous oxide (Cu_2O) + Magnesium oxide (MgO)

Wartenberg and Prophet, 1932

mol%	f.t.	mol%	f.t.
0	1240	50	1340
10	1200	60	1475
20	1190	70	1620
30	1200	100	2800
40	1260		

Cuprous oxide (Cu_2O) + Cupric oxide (CuO)

Reuter and Schröder, 1954 (fig.)

mol%	χ		
	+20°	-78°	183°
0	-0.162	-0.162	-0.162
20	+0.48	+0.35	+0.27
40	1.10	0.85	0.65
60	1.70	1.30	1.05
80	2.30	1.80	1.40
100	2.75	2.30	1.75

Next Page

mol%	log. κ	mol%	log. κ
	room temp.		
0	-4.7	75	-6.25
20	-4.9	80	-6.2
40	-5.3	100	-4.9
60	-5.9		

Cuprous oxide (Cu_2O) +Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	E	%	f.t.
0	1230	-	60	1810
6	1160	1160	80	1900
20	1450	1160	100	2050
40	1670	1160		

Beryllium oxide (BeO) + Magnesium oxide (MgO)

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	E	%	f.t.	E
0	2530	-	50	1670	1670
10	2250	-	60	1700	1670
20	1860	1670	70	1900	-
30	1780	1670	80	2150	-
40	1700	1670	100	2800	-

Ruff, Ebert and Krawczynski, 1933 (fig.)

mol%	f.t.
70	2400
40	1955 E
25	2000
0	2450

Beryllium oxide (BeO) + Calcium oxide (CaO)

Ruff, Ebert and Krawczynski, 1933 (fig.)

mol%	f.t.
80	2420
50	1650
36	1475 E
20	1700
0	2450

Beryllium oxide (BeO) + Titanium oxide (TiO_2)

Trzebiatowski, Drys and Baran, 1953

mol%	f.t.	mol%	f.t.
20	1890	60	1690
30	1835	65	1710
40	1730	70	1675 E
-	1750	80	1725
50	1700	90	1790
-	1725	-	1805
55	1725	100	1810 (1+1)

Beryllium oxide (BeO) + Zirconium oxide (ZrO_2)

Ruff, Ebert and Stephan, 1930

mol%	f.t.	mol%	f.t.
0	2400	50	2180
10	2380	55	2215
15	2180	60	2225
20	2180	65	2250
25	2180	70	2270
30	2180	75	2300
35	2180	80	2320
40	2180	85	2540
45	2180	90	2635

Ruff, Ebert and Wartenberg, 1931 (fig.)

%	f.t.	E	%	f.t.	E
53	2410	-	74	2250	-
63	2370	2220	78	2300	-
71	2220	-	80	2310	-

Wartenberg and Werth, 1930 (fig.)

%	f.t.	%	f.t.
0	2570	70	2410
20	2570	76.8	2540 (3+2)
40	2500	80	2510
50	2450	84	2070 E
64	2100	100	2700

Beryllium oxide (BeO) + Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	E	%	f.t.	E
0	2530	-	80	1890	1890
20	2220	-	90	1920	1890
40	2030	1890	100	2050	-
60	1950	1890			

Beryllium oxide (BeO) + Indium oxide (In_2O_3)

Watt and Milligan, 1953

X - ray diffraction studies.

Magnesium oxide (MgO) + Calcium oxide (CaO)

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	%	f.t.
100	2580	40	2460
70	2450	20	2540
62	2280-E	0	2800

Ruff, Ebert and Krawczynski, 1933 (fig.)

mol%	f.t.
75	2690
45	2360 E
30	2620

Rankin and Merwin, 1916 (fig.)

%	f.t.	%	f.t.
100	2570	40	2600
80	2450	20	2730
67	2300 E	0	2800

Magnesium oxide (MgO) + Strontium oxide (SrO)

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	%	f.t.
100	2430	60	2070
80	2120	20	2560
73	1935 E	0	2800

Magnesium oxide (MgO) + Barium oxide (BaO)

Wartenberg and Prophet, 1932

%	f.t.	%	f.t.
100	1923	80	1660
90	1660	65	1920
85	1500 E		

Magnesium oxide (MgO) + Zinc oxide (ZnO)

Rigamonte, 1946

mol%	lattice constant a (in Å)
0	4.203
17.7	.218
24.4	.222
24.7	.224
35.8	.230
45.3	.231
53.9	.228
67	.230 (saturated)

Magnesium oxide (MgO) + Cupric oxide (CuO)

Rigamonti, 1947

Mixed crystals up to 25 mol%. Probably a complex (3+7)

Magnesium oxide (MgO) + Manganese oxide (MnO)

Pettersson, 1946

mol%	lattice constant a (in Å)
67	4.365
55	.35
50	.34
44	.32
32	.30
25	.28

Magnesium oxide (MgO) + Ferrous oxide (FeO)

Pettersson, 1946

mol%	lattice constant a (in Å)
60	4.27
50	.265
40	.25
34	.24
20	.23

Shashkina and Guerasimov, 1953

mol%	Q mix.	mol%	Q mix.
0	0	60	-1320
1	-30.3	70	-1310
2	-60.5	80	-1160
5	-159	90	-829
10	-299	95	-533
20	-580	97	-363
30	-826	98	-270
40	-1140	99	-147
50	-1210	100	0

Magnesium oxide (MgO) + Nickel oxide (NiO)

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	tr.t.	%	f.t.	tr.t.
100	1990	-	40	2600	2400
90	2170	2050	20	2700	2550
80	2300	2150	0	2800	2800
60	2570	2310			

Magnesium oxide (MgO) + Cobalt oxide (CoO)

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	E
100	1810	-
90	2050	1950
80	2230	2080
70	2370	2190

Robin, 1955

mol%	Curie point	mol%	Curie point
0	690	72	860
26	740	100	900
51	800		

Eremenko and Beinish, 1956 (fig.)

mol%	lg R 500°	600°	700°	800°	900°
0	8.5	7.8	7.2	6.5	5.3
10	7.2	6.4	5.7	5.0	4.5
25	5.8	4.8	4.0	3.5	3.0
50	3.7	3.0	2.2	1.7	1.5
75	2.2	1.8	1.2	0.5	0
90	2.0	1.5	0.8	0	- 0.8
100	1.9	1.1	0.5	-0.2	- 1.1

Eremenko and Beinish, 1956 (fig.)

mol%	lg R annealed at 1800° 100°	200°	300°
40	207	44.3	18.6
45	-	-	13.0
50	18.9	8.4	5.30
55	3368	689	171
60	2191	370	99.4
63	40.0	8.1	331
67	61.8	14.9	4.97
70	230	32.8	10.4
75	913	134	48
80	5619	456	102
90	1633	126	125
100	18600	800	35.5

Magnesium oxide (MgO) + Titanium oxide (TiO_2)

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	E	
100	1825	-	
91	1645	1645	
80	1680	-	(1+2)
70	1625	1625	
60	1750	-	
55	1850	-	(2+1)
50	1840	-	
44	1850	1800	
40	1800	1800	
30	2000	-	
0	2800	-	

Magnesium oxide (MgO) + Zirconium oxide (ZrO_2)

Wartenberg and Werth, 1930 (fig.)

%	f.t.	%	f.t.
0	2800	79.2	2100 (1+1)
40	2300	90	1900
55	1950	92	1650
60	1800	100	2700

Ebert and Cohn, 1933

mol%	f.t.	mol%	f.t.
90	2547	50	2150
80	2448	45	2190
70	2390	40	2240
65	2346	30	2314
60	2305	20	2447
55	2193	10	2552

Zernova, 1939

mol%	f.t.	mol%	f.t.
100	2715	63	2242
96	2669	60	2169
94	2645	53	2086
91	2622	48	2086
87	2574	42	2110
83	2550	36	2145
78	2502	30	2205
76	2480	23	2288
71	2407	16	2384
67	2320	10	2491

Avgustinik and Anzelevich, 1953

Dielectric constant according to the temperature of tempering .

Magnesium oxide (MgO) + Cerium oxide (CeO_2)

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	%	f.t.
100	-	70	2330
97.5	2370	60	2370
95	2240 E	50	2450
90	2240	30	2600
80	2280	0	2800

Magnesium oxide (MgO) + Aluminum oxide (Al_2O_3)

Rankin and Merwin, 1916 (fig.)

%	f.t.	%	f.t.
0	2800	92	1925 E + I
55	2030 E	92.5-100	- E + I
71.6	2135 (1+1)	98	1925 E + I
71.6-92.5	- E	100	2050 I
71.6-97.5	- E + II		2050 II

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	E
0	2300	-
50	2050	-
55	1980	1990
60	2075	1990
70	2110	1990
72	2115	1990 (1+1)
77	2110	2000
80	2105	2050
90	2050	2020
93	2005	2005
100	2050	-

Wartenberg and Prophet, 1932

mol%	f.t.	mol%	f.t.
100	2150	50	2110 (1+1)
90	2025	40	2075
85	2000	33	1190
83	1990	30	2085
80	2020	20	2330
70	2070	10	2580
60	2110	0	2800

Magnesium oxide (MgO) + Chromium sesquioxide (Cr_2O_3)

Bielanski and Deren, 1955

κ (in mhos)	f.t. cooling	f.t. heating
	20 mol %	
32000	780	780
100000	600	720
320000	480	630
1000000	390	560
2500000	300	520
	50 mol %	
40	800	800
100	540	610
1000	310	480
10000	190	400
100000	110	400
1000000	-	330
	80 mol %	
320	810	810
1000	600	700
10000	380	520
100000	230	420
1000000	-	380

Eremenko and Beinish, 1956 (fig.)

mol%	R annealed at 1450° 100°	300°	400°	500°
0	-	-	-	1347.10^4
10	111375	3272	1337	660
25	1252	38.4	13.5	7.4
38	3718	53.6	14.0	5.8
75	5867	108	45.1	25.4
90	3709	38.3	15.5	7.5
95	3094	29.6	11.3	6.04
100	-	352	94	55.4

mol%	R annealed at 1450° 600°	700°	800°	900°
0	38180.10^4	5390.10^4	1010.10^4	235.10^4
10	415	280	219	159
25	4.50	3.17	2.51	1.95
38	3.34	2.20	1.61	1.30
75	17.9	14.0	11.9	10.2
90	5.44	4.75	4.09	3.19
95	4.27	3.34	2.88	2.56
100	36.9	29.1	24.3	19.8

mol%	R annealed at 1850° 100°	300°	400°	500°
0	-	-	-	-
10	9200	248	90	47.4
25	1110	32.2	11.7	6.2
30	1616	35.0	10.0	4.57
40	925	14.4	4.59	2.32
50	-	9371	1013	293
60	-	-	22660	4651
70	2309	86	30.0	18.4
75	1394	38.9	13.2	7.6
90	266	9.4	3.9	2.3
95	266	9.3	4.55	2.80
100	1860	125	47	33.7
	600°	700°	800°	900°
0	-	910.10^5	228.10^5	855.10^4
10	30.9	22.0	16.8	13.7
25	4.0	3.00	2.50	2.20
30	3.02	2.00	1.73	1.47
40	1.53	1.15	0.87	0.75
50	177	177	83.3	68.7
60	1798	784	287	112
70	14.2	10.7	9.6	7.9
75	6.1	4.9	4.21	3.56
90	1.76	1.55	1.47	1.38
95	2.16	1.96	1.73	1.62
100	19.0	17.8	15.9	15.0

Magnesium oxide (MgO) + Mangano-manganic oxide (Mn_3O_4)

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	%	f.t.
100	1560	40	2500
80	1800	30	2570
60	2200	0	2800

Magnesium oxide (MgO) + Ferrosoferric oxide (Fe_3O_4)

Wartenberg and Prophet, 1932 (fig)

%	f.t.	%	f.t.
100	1600	40	2450
80	1800	30	2550
60	2120	0	2800

Calcium oxide (CaO) + Barium oxide (BaO)

Grey, 1950 (fig.)

mol%	thermoionic (pulse) emission
100	0.3
80	2.2
60	3.8
40	3.4
20	2.8
0	0

Calcium oxide (CaO) + Manganous oxide (MnO)

Pettersson, 1946

mol%	lattice constant (in Å)
35	4.685
51	4.62
68	4.56

Calcium oxide (CaO) + Lead dioxide (PbO_2)

Le Chatelier, 1893

t	dissociation pressure (2+1)
880	47
940	112
950	117
1020	350
1060	557
1070	570
1100	940
1110	1040

Calcium oxide (CaO) + Titanium oxide (TiO_2)

De Vries , Roy and Osborn, 1954 (fig.)

wt%	f.t.	E		
0	2570	-		CaO
15	-	1695		"
25	-	1695		"
32	1860	1695		"
39 E	1695	1695		
42	1750	1695	(3+2)	
48	1900	1695	(1+1)	
58	1970	-		
70	1840	1460		
83 E	1460	1460		
90	1670	1460		
100	1830	-		

Calcium oxide (CaO) + Zirconium oxide (ZrO_2)

Wartenberg, Linde and Jung, 1928 (fig.)

%	f.t.	E	%	f.t.	E
0	2580	-	60	2560	2300
20	2450	2220	70	-	2400
30	2370	2220	80	-	2390
40	-	2220	100	2670	-
50	2480	2230			

Ruff, Ebert and Stephan, 1929

mol%	f.t.		mol%	f.t.
100	2687			
90	2620		46	2345
80	2575		43	2315
70	2475		40	2325
68	2430		37	2320
67	2385		34	2300
64	2370		30	2970
61	2325	(1+1)	25	2330
60	2300		20	2445
52	2325		10	2565
50	2330		0	2580

Avgustinik and Anzelevich, 1953

Dielectric constant, according to the temperature of tempering .

Hund, 1952

%	lattice constant (in Å)
97.5	5.1147
95	.1137
92.5	.1196
90	.1263
87.5	.1276
85.0	.1275
82.5	.1282

Calcium oxide (CaO) + Aluminum oxide (Al_2O_3)

Neumann, 1910

%	f.t.	%	f.t.
31.2	1650	60.3	1530
37.8	1525	64.5	1590
45.0	1410	67.5	1575
47.6	1385	70.8	1630
52.2	1400	75.2	1710
54.8	1385	100	2000

Rankin, 1915

%	f.t.	
0	2570	
37.8	-	(3+1)
41	1535	
50	1395	E
52.5	1455	(5+3)
53	1400	E
64.6	1600	(1+1)
66.5	1590	E
75.2	1720	(3+5)
76	1700	E
100	2050	

Wartenberg, Linde and Jung, 1928

%	f.t.
53	1400 E
57	1460
65	1600
100	2055

Filonenko and Lavrov, 1950

%	f.t.	%	f.t.
100.0	2050	85.0	1810
94.0	1960	83.6	1800
91.6	1910	82.5	1750
90.6	1890	80.8	1730
89.9	1870	79.4	1770
89.2	1850	78.5	1760
88.5	1850	77.5	1750
87.8	1850	76.3	1750
85.4	1830	75.0	1710

Calcium oxide (CaO) + Ferric oxide (Fe_2O_3)

Kohlmeyer and Hilpert, 1910

%	f.t.	E I	E II	E III	tr.t.
100	1563.5	-	-	-	-
98.18	1547	1426			
96.44	1521	1405	-	-	-
95.43	1526	-	-	-	-
94.16	1494	1400	1328	-	-
93.06	1470	1407	"	-	-
91.93	1467	"	-	1266	1030
90.99	1443	1397	1323	1240	-
89.52	1410	-	1313	1247	-
88.50	1386	-	"	1237	-
86.91	1340	-	1320	1242	-
83.49	1298	-	-	1237	-
82.69	1292	-	-	1246	925
81.02	1395	-	-	-	-
80.38	1364	-	-	1212	-
79.72	1338	-	-	1210	-
79.14	1305	-	-	1210	-
78.02	1298	-	-	1196	-
76.36	1277	-	-	1206	-
75.52	1237	-	-	1207	-
74.77	1226	-	-	1207	-
74.00	1202	-	-	-	-
73.62	1252	-	-	1170	-
72.43	1344	-	1197	-	-
70.93	1362	-	1202	-	-
69.96	1381	-	1207	-	-
69.10	1418	-	1200	-	-
68.10	-	-	-	-	-
67.33	1437	-	1201	-	-
65.49	1458	-	-	-	-
65.02	1439	1405	1349	1216	-
62.57	1418	1390	-	1212	-
61.55	1417	1403	-	1218	-
60.51	1455	1402	-	1230	-
58.73	-	1405	-	1216	-
57.25	1491	1414	-	1218	-
56.12	1551	1424	-	-	-
54.90	-	1430	-	-	-
52.54	-	-	-	-	-
48.78	-	1420	-	-	-
41.58	-	1410	-	-	-
33.41	-	1413	-	-	-
24.02	-	1412	-	-	-
7.74	-	1411	-	-	-

%	d	%	d
100	5.190	62.5	4.743
90.5	5.045	59	4.724
90	5.025	50	4.683
85	4.955	49	4.671
82.5	4.878	44	4.480
78	4.844	36	4.213
73	4.785	0	3.316
70	4.833		
66.7	4.723		

Hilpert and Kohlmeyer, 1909 (fig.)

mol%	f.t.	E	crystall. phase
0	-	-	A
25	-	1410	transit. (3+1)
30	-	1410	-
31	1550	1400 E	tr.t. = 1220°
35.5	-	-	
40	1455	-	(3+2)
50	-	1200 E	tr.t. = 920°
60	1400	-	(2+3)
61	-	1241 E	-
80	1470	1325	(1+3)
100	1565	1325	-

mol%	d
0	3.32
20	3.73
40	4.31
50	4.67
60	4.74
80	4.83
100	5.19

Nacken and Grunewald, 1926

mol%	f.t.	tr.t. 1	tr.t. 2
35	1412	1217	-
37	1390	"	-
39	1364	"	-
41	1355	"	-
43	1342	"	-
45	1324	1230	1212
47.5	1275	1217	-
50	1240	-	-
52.5	1235	1225	1215
55	1255	1217	-
57.5	1245	1217	-
60	1245	1200	-
0	(2570)		
100	(1565)		
(1+1)	(2+1)	(5+1)	

Konarzewski, 1931

mol%	f.t.	mol%	f.t.
30	1480	50	1220 (1+1)
33.3	1440 (2+1)	60	1190
40	1410	70	1190

Calcium oxide (CaO) + Chromium sesquioxide (Cr_2O_3)

Eremenko and Beinish, 1956.

mol%	R annealed at 1800° 100°	200°	300°
40	207	44.3	18.6
45	-	-	13.0
50	18.9	8.4	5.30
55	3368	689	171
60	2191	370	99.4
63	40.0	8.1	331
67	61.8	14.9	4.97
70	230	32.8	10.4
75	913	134	48
80	5619	456	102
90	1633	126	125
100	18600	800	35.5

mol%	R annealed at 1800° 400°	450°	500°
40	11.3	8.7	7.3
45	8.3	7.0	6.1
50	3.76	3.06	2.51
55	76.9	56.6	44.4
60	41.6	31.1	25.2
63	1.79	1.72	1.45
67	2.70	2.55	2.12
70	6.2	4.56	3.68
75	19.7	15.2	13.1
80	34.5	19.4	17.5
90	14.4	10.5	8.4
100	47	-	33.7

mol%	R annealed at 1800° 550°	600°	650°
40	6.2	5.40	4.84
45	5.32	7.77	4.23
50	2.17	1.95	1.85
55	35.0	28.8	24.0
60	20.5	17.4	15.6
63	1.28	1.11	0.98
67	1.72	1.56	1.39
70	3.12	2.77	2.52
75	10.9	9.3	8.4
80	14.3	12.1	10.3
90	7.2	6.1	5.7
100	20.8	19.1	-

mol%	R annealed at 1800° 700°	800°	900°
40	4.22	3.24	2.36
45	3.77	3.31	3.06
50	1.68	1.43	1.19
55	20.7	14.3	11.8
60	13.9	12.0	11.0
63	0.92	0.85	0.85
67	1.25	1.1[illegible]	1.19
70	2.26	1.97	1.90
75	7.3	6.2	5.9
80	9.2	7.7	6.9
90	5.1	4.67	4.18
100	17.8	15.9	15.0

Strontium oxide (SrO) + Barium oxide (BaO)

Grey, 1950 (fig.)

mol%	thermoionic (pulse) emission
100	0.3
80	1.0
60	2.4
40	4.3
33	4.6
20	2.1
0	0

Strontium oxide (SrO) + Titanium oxide (TiO_2)

Trzebiatowski and Drys, 1955

mol%	f.t.
20 - 30	1740 E
33	1800 (2+1)
50	2020 (1+1) and (3+2)
70 - 80	1440 E

Strontium oxide (SrO) + Zirconium oxide (ZrO_2)

Wartenberg and Werth, 1930

%	f.t.	%	f.t.
0	2430	75	2350
10	2390	80	2200
20	2335	82	2130 E
30	2285	90	2480
40	2215 E	100	2690

(1+1)

Strontium oxide (SrO) + Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	E	
0	2430	-	
47	2015	-	(1+1)
60	1900	1805	
70	1800	1805	
80	1920	1805	
100	2050	-	

Barium oxide (BaO) + Nickel oxide (NiO)

Lander, 1951

mol%	f.t.	
100	1650	
-	1200	E
50	1240	(1+1)
-	1080	E
25	1160	(3+1)
0	1923	

Barium oxide (BaO) + Zirconium oxide (ZrO_2)

Wartenberg and Werth, 1930 (fig.)

%	f.t.	%	f.t.
0	1910	75	2390
40	2640	85	2290
63	2600	100	2700

Barium oxide (BaO) + Titanium oxide (TiO_2)

Statton, 1951 (fig.)

mol%	m.t.	f.t.	
100	-	1725	
97.5	1625	1700	
90	1590	1630	
87.5	1588	1588	E
85	1590	1625	
80	1592	1640	(1+1)
75	1594	1625	E
70	1597	1665	
67	1598	1675	(1+2)
60	1599	1640	E
58	1600	1680	
50	1600	1700	(1+1)
45	1599	1645	E
40	1598	1675	
33	1592	1694	(2+1)
30	1590	1670	
25	1588	1645	E
20	1585	1675	
12	1580	1695	
0	-	1923	

Trzebiatowski, Drys and Berak, 1954

mol%	f.t.	mol%	f.t.
33	1820	67	1315 (1+2)
40	1650	68	1305 E
44	1585 E	75	1390
50	1610 (1+1)	80	1465 (1+4)
60	1375	82	1455 E
65	1310 E	100	1810

Barium oxide (BaO) + Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932

%	f.t.	E	
0	1925	-	
40	2000	-	(1+1)
60	1910	1875	
65	1880	1875	
70	1925	1875	
80	2000	1875	
90	2040	-	
100	2050	-	

Toropov and Galakhov, 1952

%	f.t.	%	f.t.
0	1930	40	1830 (1+1)
10	1750	55	1790
15	1660 E	70	1880
20	1750 (3+1)	80	1900 (1+6)
22.5	1710 E	83	1890 E
30	1800	100	2050

Lead oxide (PbO) + Cupric oxide (CuO)

Cunningham, 1914

%	f.t.	E	%	f.t.	E
0	875	-	11.73	774	-
1.87	850	-	14.30	746	698
3.40	838	-	21.0	727	698
5.0	818	698	26.0	712	702
6.10	808	-	34.0	-	698
7.35	797	-	45.0	712	698
10.0	784	698	100	-	-

Lead oxide (PbO) + Titanium oxide (TiO_2)

Belyaev and Nesterova, 1952

mol%	f.t.	mol%	f.t.
0.0	884	15.0	965
5.0	839	20.0	1030
10.0	880	25.0	1099

Lead oxide (PbO) + Stannic oxide (SnO_2)

Urazov, Speranskaya and Gulyanitskaya, 1956.

mol%	f.t.	E	mol%	f.t.	E
0	865	-	12.0	1010	-
1.8	855	-	14.0	1060	-
2.5	850	-	15.0	1060	-
3.8	897	-	20.0	1100	-
5.0	905	855	25.0	-	-
7.0	930	855-848	33.4	1280	-
8.5	970	855	40	1290	-
10.0	990	840			

Lead oxide (PbO) + Ferric oxide (Fe_2O_3)

Kohlmeyer, 1913

%	f.t.	tr.t. I	II	III	IV	V
17	850	-	-	-	815	-
19.26	903	-	-	-	825	-
20	925	-	-	-	815	-
21.23	937	-	-	-	827	-
25	1060	-	945	-	-	850
26.35	1077	-	920	1010	830	-
27	-	-	-	-	-	-
30	1137	-	940	-	-	845
32.31	1175	1140	927	1030	830	-
35	1197	1135	-	990	-	850
37	1212	1125	-	997	-	857
36.7	1220	-	-	1030	-	-
40	1227	1138	927	1017-997	-	850
41.72	1235	1135	935	995	-	830

%	f.t.	tr.t. VI	VII	VIII	IX	XV
0	885	-	-	640	560	-
5	810	743	-	-	-	-
7.5	785	750	-	-	-	-
10	762	-	-	-	570	-
12	755	755	-	-	570	-
14.16	795	757	-	-	-	-
15.17	827	753	-	653	-	-
17	850	763	-	-	-	-
19.26	903	753-725	-	660	-	-
20	925	-	765	-	-	-
21.23	937	-	765	-	550	-
25	1060	-	765	-	-	-
26.35	1077	-	770-730	-	-	-
27	1137	-	-	-	-	-
30	1175	-	765	660	565	-
32.31	1197	-	775	650	-	-
35	1212	-	760	-	-	-
37	1212	-	770	650	-	-
36.7	1220	-	760	-	-	-
40	1227	-	-	640	-	785
41.72	1235	-	-	650	-	775

%	f.t.	tr.t. I	III	V	VIII	X
43.68	1300	1160	995	835	-	-
48.20	1340	-	-	860	-	-
48.68	1352	-	-	850	-	-
57.07	1412	-	-	850	653	-
58.88	1462	-	-	840	-	-
62.55	1485	-	-	-	-	-
68.26	1510	-	-	-	-	-
74.12	1535	-	-	840	-	-
80.23	1535	-	-	840	610	1435
86.56	1590	-	-	-	-	1460
100	1565	-	-	-	-	1520-1470

%	f.t.	tr.t. XI	XII	XIII	XIV	XV
43.68	1300	-	1230	1112	960	780
48.20	1340	-	1230	1110	970	780
48.68	1352	-	1210	1110	950	780
57.07	1412	1352	1275-1195	" -1070	60	-
58.88	1425	1360	1235	1075	60	-
62.55	1462	1350	1260	1130-1075	950	-
68.26	1485	1395-1320	23	1070	950	780
74.12	1510	1380	1240	-	920	783
80.23	1535	1365	1240	1085	960	-
86.56	1590	1355	1240	1080	952	-
100	1565	1375	-	-	-	-

%	d	%	d
0	9.514	51.47	6.705
5	9.150	59.88	6.337
19.26	8.531	63.60	6.119
25	8.170	71.74	5.974
30	7.745	77.66	5.793
39.78	7.296	81.61	5.661
41.72	7.194	90.29	5.353
48.52	6.936	100	5.190
50.68	6.712		

Lead oxide (PbO) + Chromium trioxide (CrO_3)

Jaeger and Germs, 1921

mol%	f.t.	E	tr.t.	
0	879	-	-	
5.7	840	784	-	
9.1	801	786	-	
13.1	813	785	-	
14.6	815	785	-	
16.7	830	815	-	
18	838	815	744	
20	846	815	745	(6+1)
21.3	815	-	-	
22.2	854	754	-	
23.1	850	841	754	(4+1)
25.1	865	841	754	
27.3	896	840	754	
28.7	905	841	754	
30.6	913	839	754	
33.4	920	-	-	
37.6	914	785	700	(2+1)
40.2	820	786	701	
41.2	822	785	701	
43	820	786	701	
46	820	785	700	
48	820	785	700	
49	819	784	699	
50	814	783	707	

Lead oxide (PbO) + Molybdenum trioxide (MoO_3)

Jaeger and Germs, 1921

mol%	f.t.	E	mol%	f.t.	E
0	879	-	35.5	944	932
4.8	855	751	37.6	933	-
9	815	761	39.4	972	933
13.1	841	762	41	998	933
17	879	762	44.5	1031	933
23	921	761	48	1052	932
29	939	757	50	1065	-
33.4	951	-			

Lead oxide (PbO) + Tungsten trioxide (WO_3)

Jaeger and Germs, 1921

mol%	f.t.	E	mol%	f.t.	E	tr.t.
0	879	-	35.5	915	899	891
9	815	721	37.6	981	899	892
13.1	770	722	41	1038	899	892
17	780	722	46	1087	891	877
23	855	722	49	890	877	-
29	886	722	50	1123	877	-
33.4	899	-				

Lead oxide (PbO) + Vanadium pentoxide (V_2O_5)

Amadori, 1913

%	f.t.	E	min.
100	660	-	-
90	654	465	-
80	642	475	-
70	626	475	-
60	582	475	-
50	526	480	-
45	480	480	-
52	554	476	-
40	606	470	-
38	645	468	-
35	684	468	-
33	704	465	-
30	718		-
29.63	722	-	-
27	824	718	140
25	900	715	70
23	942	710	20
21.43	952	-	-
20.36	946	746	30
19.5	936	746	40
18	906	746	90
17.93	886	748	100
15	820	750	140
12	770	-	120
10	790	755	20
18.28	794	760	-
8	788	755	30
7	775	760	100
5	800	760	120
3	838	758	70
0	890	-	-

(1+1) 480° (2+1) 722° (3+1) 952°
(8+1) 794°

Belyaev and Nesterova, 1952

mol%	f.t.	mol%	f.t.
0.0	884	25.0	952
1.6	872	26.9	933
3.5	855	29.0	851
5.1	815	33.3	772
6.0	793	38.2	692
7.7	765	43.5	607
10.0	765	50.0	497
12.5	765	57.2	518
15.2	793	65.4	582
18.2	748	75.0	651
21.4	884	86.4	666
23.2	924	100	675

(8+1) (3+1) (2+1)

Lead oxide (PbO) + Bismuth trioxide (Bi_2O_3)
Belladen, 1922

%	f.t.	tr.t.	min.	E	min.
100	817	-	-	-	-
94.93	785	675	75	-	-
91.19	765	680	175	-	-
90.00	750	685	180	-	-
89.27	742	686	230	-	-
86.80	722	690	85	678	20
82.92	705	695	15	685	45
86.52	695	-	-	680	90
80.00	687	-	-	"	105
78.80	-	-	-	"	180
77.21	685	-	-	"	65
75.73	686	-	-	-	-
73.09	680	-	-	-	-
70.00	660	595	10	-	-
67.54	655	615	20	-	-
60.00	630	605	55	-	-
56.04	625	610	110	-	-
52.84	615	"	80	-	-
50.98	625	-	-	-	-
50.00	620	-	-	-	-
47.13	605	580	35	-	-
44.72	602	"	140	-	-
40.96	630	"	160	-	-
34.21	680	"	110	-	-
26.85	730	"	55	-	-
18.97	770	"	54	-	-
9.87	820	"	30	-	-
0	870	-	-	-	-

(2+1) (1+2)

Zinc oxide (Zn O) + Cadmium oxide (CdO)

Rigamonti, 1946

mol%	lattice constants (in Å) a	c
0	3.248	5.203
2.36	.255	.215
2.36	.257	.217
4.56	.263	.228
4.78	.263	.228
6.82	.265	.230
7.94	.264	.226
13.6	.265	.229
20.0	.264	.230
28.1	.265	.230

Zinc oxide (ZnO) + Cupric oxide (CuO)

Rigamonti, 1947

No mixed crystals (Roentgenograms)

Zinc oxide (ZnO) + Manganese oxide (MnO)

Rigamonti, 1946

mol%	lattice constants (in Å) a	c
0	3.248	5.203
5	.257	.218
10	.261	.225
15	.272	.241
20	.282	.258
30	.283	.259
40	.278	.250
50	.283	.259
78	.281	.256 (sat.)

Zinc oxide (ZnO) + Nickel oxide (NiO)

Rigamonti, 1946

mol%	lattice constant a (in Å)
100	4.168
90	.178
85	.181
80	.189
70	.202
60	.205
50	.205
40	.206
30	.206
35	.205 sat.

Zinc oxide (ZnO) + Cobalt oxide (CoO)

Robin, 1955

qualitative diagram of the freezing curve

Hedvall, 1913

mol%	d (mixed crystals)
	17°
41.4	5.54
28.4	.42
18.9	.41
14.8	.46
13.6	.44

Rigamonti, 1946

mol%	mixed crystals
30	hexagonal, green a = 3.25 Å c= 5.20 Å
70	cubic, rose a = 4.26 Å

Zinc oxide (ZnO) + Zirconium oxide (ZrO_2)

Wartenberg and Gurr, 1931 (fig.)

%	f.t.	%	f.t.
0	1960	40	1810
16	1820	50	1812
30	1810	58	1830
	E	100	2700

Zinc oxide (ZnO) + Aluminum oxide (Al_2O_3)

Bunting, 1932 (fig.)

mol%	f.t.		mol%	f.t.
100	2045		40	1940
90	1975		30	1880
80	1950		20	1750
70	1945		17	1720 E
60	1945	L_1+L_2	10	1830
50	1945		0	1975

Wartenberg and Reusch, 1932 (fig.)

%	f.t.
0	1970
11	1815 E
20	1860
40	1925

Zinc oxide (ZnO) + Ferric oxide (Fe_2O_3)

Toropov and Borisenko, 1952

Roentgenographic study of mixed crystals .

Zinc oxide (ZnO) + Indium sesquioxide (In_2O_3)

Chalyi and Rozhenko, 1956

%	X-ray systeme
40.55	Mixture of
46.02	solide
53.20	solution
63.03	and ZnO
77.32	
87.21	
91.10	solide
93.17	solution
94.46	of In_2O_3

Cadmium oxide (CdO) + Bismuth trioxide (Bi_2O_3)

Hauffe and Peters, 1952 (fig.)

mol%	κ 300°	κ 550°	mol%	κ 300°	κ 500°
95	0.60	10	57	12	63
90	.40	8	48	80	250
85	.42	6.3	30	6300	16000
80	.16	3.2	23	3200	6300
72	.25	5.0	10	10000	20000
67	1.60	12.6	5	16000	25000
61	10	14.0	0	120000	200000

Cupric oxide (CuO) + Nickel oxide (NiO)

Rigamonti, 1947

Mixed crystals from 100 to 65 mol% NiO (X - rays)

Manganese oxide (MnO) + Ferrous oxide (FeO)

Andrew, Maddocks and Howat, 1931

%	f.t.	%	f.t.
100	1410	50	1480 - 1495
90	1420 - 1440	40	1500 - 1515
85	1425 - 1445	35	1507 - 1525
75	1450 - 1462	30	1515 - 1535
70	1465 - 1450	20	1535 - 1545
60	1460 - 1475	10	1565 - 1580
55	1470 - 1490	0	1585

Petterson, 1946

mol%	a (Å)	mol%	a (Å)
75	4.34	50	4.385
66	.355	40	.39
60	.37	25	.41

Manganese oxide (MnO) + Titanium oxide (TiO_2)

Cain, 1920

%	f.t.	%	f.t.	
95	1320	45	1360	
90	1160	40	1420	
85	1250	35	1425	
80	1200	30	1457	(complexes)
75	1250	25	1420	
70	1180	20	1455	
65	1160	15	1485	
60	1200	10	1500	
55	1300	5	1500	
50	1280			

Lepinskikh Esin and Shavrin, 1956.

mol%	κ 1480°	κ 1470°	κ 1460°	κ 1450°
40.4	-	56000	-	48000
43.5	-	39000	-	32000
50.5	-	30000	-	25000
56.0	28000	25000	25000	23000
60.1	25000	23000	--	21000
66.0	22000	21000	-	17000

mol%	1440°	1430°	1420°	1410°	1400°
40.4	-	46000	41000	-	36000
43.5	-	28000	-	25000	-
50.5	23000	-	-	21000	
56.0	21000	-	21000	-	19000
60.1	20000	-	19000	-	18000
66.0	16000	-	16000	-	14000

mol%	1390°	1380°	1370°	1360°
40.4	-	34000	-	33000
43.5	24000	23000	-	21000
50.5	19000	19000	18000	-
56.0	18000	18000	-	17000
60.1	-	17000	-	-
66.0	-	-	12000	-

mol%	1350°	1340°	1330°	1320°
40.4	-	32000	-	31000
43.5	-	20000	-	19000
50.5	17000	-	16000	-
56.0	-	16000	-	-
60.1	14000	-	14000	14000
66.0	-	12000	-	-

mol%	1310°	1300°	1290°	1280°
40.4	-	30000	-	30000
43.5	-	18000	-	17000
50.5	16000	16000	8000	-
56.0	-	14000	-	-
60.1	-	12000	-	-
66.0	-	-	-	9000

Ferrous oxide (FeO) + Ferric oxide (Fe_2O_3)

Popel and Esin, 1956 (fig.)

mol%	σ	mol%	σ
		25° above f.t.	
3	590	7	510
4	575	8	480
5	550	8.5	480
6	525		

Nickel oxide (NiO) + Cobalt oxide (CoO)

Robin, 1955

mol%	Curie point	mol%	Curie point
0	540	77	840
13	620	94	890
50	770	100	900

Hedvall, 1915

%	d		%	d	
	1300°	1100°		1300°	1100°
0.00	-	7.45	66.78	6.63	7.15
9.12	7.28	-	83.39	6.90	-
16.70	7.15	-	90.59	6.84	6.49
33.42	6.81	7.07	100.00	6.38	6.37
50.10	6.97	6.82			

Nickel oxide (NiO) + Titanium oxide (TiO_2)

Eremenko and Beinish, 1956 (fig.)

mol%	lg R			
	650°	750°	800°	900°
100	4.8	4.0	3.8	3.2
90	4.9	4.1	3.9	3.1
80	4.8	3.8	3.5	3.0
70	4.6	3.7	3.4	2.8
60	5.2	4.6	4.2	3.4
50	5.3	4.7	4.2	3.5
40	4.0	3.5	3.3	2.8
30	3.0	2.5	2.3	-
20	3.0	2.4	2.2	-
10	2.3	2.0	1.8	1.5
0	2.7	2.2	2.0	1.8

Nickel oxide (NiO) + Zirconium oxide (ZrO_2)

Wartenberg and Gurr, 1931 (fig.)

%	f.t.	%	f.t.
0	2090	60	2100
20	2070	75	2280
40	2050 E	100	2700
50	2070		

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	%	f.t.
0	1990	60	2000
20	1980	70	2100
40	1960	100	2700

Nickel oxide (NiO) + Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	E	
0	1990 corr.	-	
10	1950	1875	
12	1870	1875	
20	1920	1875	
40	1990	-	
57.5	2020	-	(1+1)
70	2000	-	
80	1975	1955	
85	1950	1955	
90	1930	1955	
100	2050	-	

N.B. corrects results of Wartenberg and Gurr,1931

Nickel oxide (NiO) + Tungsten oxide (WO_3)

van Liempt, 1925

mol%	f.t.	E	tr.t.
50	705	-	588 and 577
51.2	685	-	556
52.6	671	612	547
54.0	653	630	515
55.5	628	628	493
57.1	643	630	-
58.8	678	633	483 (1+1)
60.6	697	632	-
62.4	717	623	-
64.5	727	629	-
66.6	738	-	547
71.4	731	-	-

Hoermann, 1928

%	f.t.	E	min.	tr.t. (1+1)	min.
79.0	700	-	-	-	590-583
81.1	-	-	-	-	547
82.15	652	628	175	-	515
83.2	-	-	-	-	493
84.9	665	629	220	-	-
85.3	677	629	180	-	475
86.3	696	626	125	-	440
87.4	-	-	-	-	-
89.5	738	-	-	-	-
90.5	728	728	245	-	-
91.6	755	730	175	-	-
92.0	768	728	160	-	-
				(1+2)	
93.0	824	727	125	784	56
93.7	-	727	95	780	90
94.7	-	727	72	777	180
95.1	-	-	-	776	186
95.8	-	-	-	773	180
96.4	-	-	-	773	70
96.5	-	-	-	722	60

Cobalt oxide (CoO) + Titanium oxide (TiO_2)

Eremenko and Beinish, 1956 (fig.)

mol%	lg R 600°	700°	800°	900°
0	0	- 0.4	- 0.8	- 1.5
10	0.3	- 0.1	- 0.5	- 1.0
30	1.0	- 0.5	- 0.4	0.3
40	1.5	- 1.1	- 0.8	0.4
50	3.8	2.8	2.5	2.2
60	3.2	2.8	2.4	2.2
70	3.4	2.7	2.4	2.2
80	3.7	3.0	2.7	2.5
90	4.3	3.7	3.2	2.8
100	5.0	4.4	3.8	3.4
	500°			
90	4.95			

Cobalt oxide (CoO) + Zirconium oxide (ZrO_2)

Wartenberg and Gurr, 1931 (fig.)

%	f.t.	%	f.t.
0	1935	58	1850
20	1850	70	1890
40	1850 E	100	2700

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	%	f.t.
0	1810	60	1750
20	1710	70	1850
40	1720	100	2700

Cobalt oxide (CoO) + Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	E
0	1930	-
10	1775	1715
13	1715	1715
20	1765	1715
40	1900	-
57.5	1960	- (1+1)
70	1950	-
80	1925	1910
85	1910	1910
90	1950	1910
100	2050	-

Cobalt oxide (CoO) + Ferric oxide (Fe_2O_3)

Robin, 1955 (fig.)

mol%	Curie point	mol%	Curie point
74.5	1000	55.5	520
71.8	880	55.5	670
71.8	520	65.2	820
		77.0	830
		91.8	810
		93.5	720
		95.1	520

Titanium oxide (TiO_2) + Zirconium oxide (ZrO_2)

Wartenberg and Gurr, 1931 (fig.)

%	f.t.	%	f.t.
0	1850	40	1800
10	1800	60	1900
20	1760	80	2100
22	1750 E	100	2700

Coughanour, Roth and De Prosse, 1954

mol%	wt%	m.t.	f.t.	
90	93.28	-	-	
75	82.23	1900	-	
66.67	75.51	1845	-	
60	69.81	1825	-	
55.56	65.84	1820	-	
52.94	63.44	1825	1870	
50	60.71	1820	1860	
47.00	57.81	1790	1835	
44.44	55.23	1775	1820	
40	50.69	1760	1825	
33.35	53.54	1755	1800	
25	33.95	1765	1780	
20	27.83	1755	1760	
16.67	23.57	1760	1780	
10	14.63	1785	1805	
5	7.51	1810	1830	
0	0	-	1840	(1+1)

Eremenko and Beinish, 1956 .

mol%	R . 10_4 annealed at 1350°		
	500°	550°	600°
100	92	33.1	16.5
90	112	-	38.5
80	264	-	59.0
70	-	220	103
60	-	-	175
50	-	-	184
40	-	293	133
30	-	229	75.4
20	-	121	42.2
10	-	94.2	20.6
0	-	57.3	18.2

mol%	700°	800°	900°
100	5.86	2.5.	1.22.
90	13.8	3.45.	1.29.
80	59.0	4.21.	1.38.
70	103	7.99.	2.42.
60	175	12.2	3.62.
50	184	12.9	4.02.
40	133	8.27.	2.70.
30	75.4	4.99.	1.62.
20	42.2	2.08.	6.82.
10	20.6	1.96.	0.620
0	18.2	0.770	0.276

Titanium oxide (TiO_2) + Stannic oxide (SnO_2)

Lietz and Nodop, 1955 (fig.)

mol%	a (in Å)	
	1	2
	heated at 900 - 1000°	
100	4.735	-
80	.710	-
60	.680	-
40	.695	4.615
20	.710	.615
0	-	.590

a = lattice constant

Titanium oxide (TiO_2) + Titanium sesquioxide (Ti_2O_3)

Junker, 1936

%	f.t.
0	1800
17	1550 E
40	1630
60	1760
100	1900

Titanium oxide (TiO_2) + Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	E	
0	1850	-	
10	1760	1700	
20	1700	1700	
30	1840	-	
40	1890	-	
50	1900	1870	(1+2)
60	1875	1870	
70	1905	1870	
80	1950	-	
90	2020	-	
100	2050	-	

Bunting, 1933

%	f.t.	%	f.t.	%	f.t.
10	1760	40	1830	70	1880
15	1730	50	1855	75	1905
20	1715 E	56	1860 (1+1)	85	1965
30	1775	60	1855	100	2045
35	1810	62	1850 E		

Titanium oxide (TiO_2) + Ferrosoferric oxide (Fe_3O_4)

Junker, 1936

%	f.t.	%	f.t.
0	1800	50	1500 E
20	1650	60	1550
40	1550		

Titanium oxide (TiO_2) + Niobium pentoxide (Nb_2O_5)

Roth and Coughanour, 1955

mol%	f.t.	m.t.	mol%	f.t.	m.t.
5	1800	1780	67	1475	1465
10	-	1510	75	1480	1480
14.29	-	1490	80	1470	1465
20	-	1475	83.33	1475	-
25	-	1475	85.71	-	-
33	1500	1475	90	1485	1475
40	1430	1470	95	1495	1485
50	1490	1490	100	1500	1500
60	1470	1465		(1+1)	(1+3)

Zirconium oxide (ZrO_2) + Cerium oxide (CeO_2)

Wartenberg and Gurr, 1931 (fig.)

mol%	f.t.	mol%	f.t.
39	2500	60	2420
48	2410	68	2540
52	2400 E		

Zirconium oxide (ZrO_2) + Thorium oxide (ThO_2)

Ruff, Ebert and Woitinck, 1929

mol %	m.t.	f.t.	mol %	m.t.	f.t.
0	2700	2700	50	2840	2975
10	2705	2745	60	2885	3005
20	2720	2800	70	2940	3025
25	2755	2850	80	2995	3040
33	2795	2895	90	3050	3050
40	2795	2940			

Zirconium oxide (ZrO_2) + Niobium pentoxide (Nb_2O_5)

Roth and Coughanour, 1955

mol%	f.t.	m.t.	mol%	f.t.	m.t.
5	-	1850	50	1600	1440
10	-	1725	67	1480	1435
12.50	-	1675	75	1440	1435
14.29	-	1670	80	1450	1435
16.67	-	1650	83.33	1460	1435
20	-	1625	85.71	1465	1445
25	-	1600	90	1475	1455
33	-	1550	95	1490	1475
40	1670	1460			

Zirconium oxide (ZrO_2) + Aluminum oxide (Al_2O_3)

Wartenberg, Linde and Jung, 1928

%	f.t.	E	%	f.t.	E
100	2055	-	40	-	1110
90	1170	1110	20	-	1110
80	-	1110	0	2670	-
60	-	1110	$L_1 + L_2$		

Zirconium oxide (ZrO_2) + Ferric oxide (Fe_2O_3)

Wartenberg and Gurr, 1931 (fig.)

mol%	f.t.	mol%	f.t.
100	1600	40	1630
80	1530	25	1970
70	1510	0	2700
60	1520		

Zirconium oxide (ZrO_2) + Chromium sesquioxide (Cr_2O_3)

Wartenberg and Werth, 1930

mol%	f.t.
65	2340
55	2320
40	2360
25	2440

Zirconium oxide (ZrO_2) + Yttrium oxide (Y_2O_3)

Hund, 1951

mol%	d	mol%	d
0	5.86	35.72	5.65
8.47	.92	45.44	.55
17.24	.78	55.54	.47
26.32	.75		

mol%	lattice constant (in Å)
0	-
7.82	5.123
16.03	.127
24.66	.136
33.74	.157
43.28	.183
53.38	.212
64.38	.228
75.36	.231
87.30	.231

Zirconium oxide (ZrO_2) + Mangane oxide (Mn_3O_4)

Wartenberg and Gurr, 1931 (fig.)

mol%	f.t.	mol%	f.t.
100	1700	40	1650
80	1650	20	1950
60	1610	0	2700

Wartenberg and Reusch, 1932 (fig.)

mol%	f.t.	mol%	f.t.
100	1560	40	1550
80	1500	30	1670
60	1470	0	2700

Cerium oxide (CeO_2) + Uranium dioxide (UO_2)

Magneli and Kihlborg, 1951 (fig.)

mol%	lattice constant (in Å)
0	5.41
20	.42
40	.43
60	.44
80	.453
100	.47

Hund, Wagner and Peetz, 1952

mol%	d	lattice constant (in Å)
	25°	
10.1	-	5.421
20.2	7.99	.416
30.3	-	.421
40.3	8.53	.427
50.4	8.87	.433
60.3	8.50	.440
100	10.8	5.458

Rüdorff and Valet, 1953

mol%	lattice constant (in Å)
0	5.411
18	.421
20	.423
24	.424
33.4	.430
34	.430
46	.437
49	.439
50	.440
59.5	.445
65	.448
68	.450
79	.457
83	.459
95	.466
100	.469

Cerium oxide (CeO_2) + Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	%	f.t.
0	2600	60	1825
20	1850	80	1900
40	1785 E	100	2050
50	1800		

Cerium oxide (CeO_2) + Cerium sesquioxide (Ce_2O_3)

Brauer and Gradinger, 1954

mol%	lattice constant (in Å)	
0	5.400	phase I
2.7	.399	(CeO_2)
8.2	.401	"
9.7	.400	"
13.7	.483	phase II
17.6	.485	mixed crystals
20.8	.487	"
23.9	.485	"
24.4	.486	"
42.8	.536	"
47.4	.542	"
49.0	.545	"
49.7	.548	"

Cerium oxide (CeO_2) + Yttrium oxide (Y_2O_3)

Mc Cullough and Britton, 1952

atom%	lattice constant (in Å)
0	5.411
9.99	.411
25.0	.405
39.9	.395
60.0	(5.374)
75.0	(5.355)
90.0	(5.324)
100	(5.304)

Values in brackets are 1/2 of the real constant for the "c" form

Cerium oxide (CeO_2) + Samarium oxide (Sm_2O_3)

Mc Cullough and Britton, 1952

atom%	lattice constant (in Å)
0	5.411
10.0	.423
20.0	.433
34.9	.441
49.9	.453
64.8	.462
80.0	(5.466)
90.0	(5.466)
100	(5.461)

Values in brackets are 1/2 of the real constant for the "c" form

Cerium oxide (CeO_2) + Lanthanum oxide (La_2O_3)

Croatto and Mayer, 1943

t	κ 0 mol%	2mol%	5mol%
300	0.0262	0.0813	0.1346
400	.0578	0.523	1.000
500	.101	2.890	4.805
600	.448	10.98	19.48
700	1.952	26.87	49.75
800	6.52	57.4	100.00
900	18.70	96.1	170.00
1000	53.79	156.1	239.00
1100	147.90	263.5	354.50
1200	340.20	416.6	561.50
1300	675.8	793.5	917.5
	15mol%	30mol%	100mol%
300	0.2512	0.1907	-
400	1.733	1.228	-
500	4.948	4.820	0.006317
600	34.70	23.96	.009133
700	82.62	64.60	.01481
800	173.80	123.7	.0660
900	317.1	263.1	.2462
1000	512.6	446.5	.8925
1100	775.0	662.0	2.821
1200	1000.0	892.7	8.520
1300	1176.0	1098	24.020

mol%	d
0	7.18
2	7.13
5	7.05
15	6.81
30	6.51
100	6.56

Cerium oxide (CeO_2) + Gadolinium oxide (Gd_2O_3)

Mc Cullough and Britton, 1952

atom%	lattice constant (in Å)
0	5.411
9.92	.415
23.1	.423
42.8	.432
67.0	(5.428)
83.1	(5.420)
100	(5.407)

Values in brackets are 1/2 of the real constant for the "c" form fluoride type

Molybdenum oxide (MoO_3) + Tungsten oxide (WO_3)

Rieck, 1943

mol%	f.t.	mol%	f.t.
0	795	15	855
1	780	16	860
2	770	21-23	905-920
3	777	27	950
6	800	30	975
8	835		

Praseodymium oxide (PrO_2) + Praseodymium sesquioxide (Pr_2O_3)

Brauer and Gradinger, 1954

mol%	lattice constant (in Å)	
1.6	5.383	phase I
3.8	5.383	"
8.4	5.457	phase II
8.9	5.456	mixed crystals
14.6	5.455	"
20.3	5.458	"
22.1	5.463	"
31.3	5.483	"
37.3	5.491	"

Praseodymium oxide (PrO_2) + Yttrium oxide (Y_2O_3)

Mc Cullough and Britton, 1952

mol %	lattice constant (in Å) a
0	5.394
14.9	5.398
30.0	5.398
50.0	(5.385)
70.0	(5.355)
85.1	(5.330)
100	(5.304)

Values in brackets are 1/2 of the constant for the "c" form .

Thorium oxide (ThO_2) + Uranium oxide (UO_2)

Dawson and Lister, 1952

t	χ	t	χ
100%		90%	
-183	14.77	-76	9.81
-75.5	10.90	+20	7.90
+20	8.74	64.5	7.22
202	7.55	129	6.45
150	7.00	180	5.57
203	6.57		
252	6.04		
290	5.73		
80%		70%	
-76	9.07	-76	7.98
+20	7.25	+20	6.35
65	6.59	61	5.87
125	5.95	28	5.19
235.5	5.47		
60%		50%	
-76	7.23	-76	6.11
+20	5.62	+20	4.75
57.5	5.16	62.5	4.30
79.5	4.95	125	3.94
120	4.58	177	3.50
144	4.39		
40%		30%	
-76	5.20	-76	4.13
+20	3.92	+20	3.05
54	3.64	64	2.72
80	3.41	125	2.38
124	3.19	132.5	2.15
144	3.01		
160	2.94		
20%		10%	
-76	2.94	+20	1.05
20	2.14		
65	1.89		
127	1.64		
179	1.48		

Dawson and Roberts, 1956

mol%	U valency	χ (by atom U)
	27°	
14.8	4.00	3450
	4.68	1840
	4.91	1280
	5.19	817
	5.46	638
27.9	4.00	2880
	4.45	2130
	4.99	1110
	5.37	667
43.1	4.00	2730
	4.32	2220
	5.02	960
	5.20	680

Rüdorff and Valet, 1953

mol%	lattice constant (in Å)
0	5.597
81	.483
89	.482
95	.475
100	.469

Thorium oxide (ThO_2) + Plutonium oxide (PuO_2)

Dawson, 1952

t	χ	t	χ
100%		75%	
-183	4.859	-183	4.644
-73	3.258	-72.5	2.771
+27	2.693	+27	2.168
60.5	2.577	61.5	2.055
74	2.543	99	1.942
133	2.426	138.5	1.852
157	2.414	173.5	1.793
		248	.754
		319	.675
50.0%		35.3%	
-183	3.426	-183	2.475
-74	2.034	-73	1.548
+27	1.614	+27	.294
59	1.490	55.5	.158
99	1.436	98	.117
142	1.410	138	.062
172	1.374	164	.034
12.1%		8.82%	
-183	1.205	+27	0.280
-72	0.552		
+26	0.344	8.72%	
61.5	0.303		
102	0.260	+27	0.352
141	0.230		
189	0.194	4.31%	
		+27	0.102
		2.19%	
		+27	0.056

Thorium oxide (ThO_2) + Aluminum oxide (Al_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	E	%	f.t.	E
0	3050	-	60	1920	1920
10	2110	1920	70	1940	"
20	2000	"	80	1950	"
30	1950	"	90	2000	"
40	1930	"	100	2050	"
50	1925	"			

Thorium oxide (ThO_2) + Yttrium oxide (Y_2O_3)

Hund and Mezger, 1952 (fig.)

mol%	d
0	9.70
15.0	.21
30.0	8.56

mol%	lattice constant (in Å)
0	5.595
20	.575
30	.565
45	.569
60	.569
90	.569
100	.569

Thorium oxide (ThO_2) + Lanthanum oxide (La_2O_3)

Hund and Dürrwächter, 1951

mol%	d	lattice constant (in Å)
	25°	
0	9.96	5.592
15.2	9.20	.611
28.8	8.52	.620
41.0	8.20	.635
51.9	7.79	.647
61.9	-	.645
70.9	-	.646

Hund, 1953

mol%	$B.10^{16}$
0	0.45
14.3	0.75
30.4	1.52

B = " Gitterstörung " (in cm^2)

Thorium oxide (ThO_2) + Uranium oxide (U_3O_8)

Hund and Niessen, 1952

mol%	d	lattice constant (in Å)
0	9.71	5.5952
1.16	9.71	.5902
2.5	9.78	.5839
4.2	10.07	.5749
6.5	10.06	.5666
9.4	9.64	.5528
13.5	9.99	.5372
17.2	10.46	.5240
23.8	10.90	.5016
31.9	-	.4852
42.2	-	.4852
55.6	-	.4853
73.7	-	.4846

Aluminum oxide (Al_2O_3) + Chromium sesquioxide (Cr_2O_3)

Bunting, 1931

mol%	m.t.	f.t.	mol%	m.t.	f.t.
0	2.045	2.045	60	2.170	-
10	.060	.075	70	.190	-
20	.080	.110	80	.205	-
30	.100	.140	90	.225	-
40	.120	.170	100	.275	2.275
50	.150	-			

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	%	f.t.
100	2275	40	2200
80	2270	20	2100
60	2250	0	2050

Cirilli, 1947

%	χ 600°	χ 850°	χ 1000°
-	-0.38	-0.35	-0.29
15	0.9	1.1	2.0
25	3.9	4.0	6.1
33	11.8	11.9	12.8
50	15.1	15.0	15.7
67	18.0	18.1	19.0
75	20.1	20.7	21.8
85	22.2	21.9	23.3
100	27.2	25.7	25.7

Selwood, Lyon and Ellis, 1951

t	χ	t	χ
	2.6 %		
27	- 1.7	- 89	- 3.1
-43	- 2.3	-125	- 4.0
-58	- 2.5	-145	- 4.2
-73	- 2.6	-189	- 6.8
	5.3 %		
27	- 3.0	-145	- 7.9
-43	- 4.4	-191	-12.3
-89	- 5.5		
	8.2 %		
27	- 6.1	- 89	- 8.6
-43	- 6.9	-145	-12.1
-58	- 7.6	-189	-17.5
-73	- 8.0		
	17.6 %		
27	-10.9	- 89	-16.1
-43	-13.4	-106	-17.3
-58	-14.5	-145	-22.0
-73	-15.2	-198	-27.9
	25.0 %		
27	-12.7	- 89	-18.5
-43	-15.9	-145	-23.6
-58	-17.2	-189	-33.8
-73	-17.8		
	33.0 %		
27	-15.0	- 89	-21.5
-43	-18.4	-106	-23.0
-58	-19.3	-145	-30.0
-73	-20.1		
	43.0 %		
27	-16.8	- 89	-22.8
-43	-20.4	-106	-24.3
-58	-20.4	-145	-34.2
-73	-21.8	-190	-52.0
	54.0 %		
27	-19.1	- 89	-26.5
-43	-22.7	-145	-36.6
-58	-23.5	-190	-47.0
-73	-24.2		
	74.6 %		
27	-22.5	- 89	-26.8
-43	-25.7	-145	-27.2
-58	-26.2	-183	-32.0
-73	-26.3		
	81.6 %		
27	-23.6	- 73	-27.2
-43	-26.3	- 89	-27.6
-58	-27.0	-133	-29.5

Eremenko and Beinish, 1956

Electric specific resistance R

mol%	100°	300°	400°
100	$1.86.10^4$	125	47
80	$2.04.10^4$	260	97
70	$12.7.10^4$	825	240
60	$15.4.10^4$	960	310
50	-	3450	1040
40	306.10^4	$24.5.10^4$	4800
20	-	48.10^4	$30.6.10^4$
10	-	-	260.10^4
0	-	-	4370.10^4

mol%	R 500°	600°	700°
100	33.7	19.0	17.8
80	56.5	41.1	33.1
70	118	79.8	62.5
60	168	111	85
50	500	322	238
40	1860	970	637
20	$50.2.10^4$	$15.4.10^4$	6930
10	$28.4.10^4$	$56.7.10^4$	$15.3.10^5$
0	-	-	24.10^6

mol%	R 800°	900°
100	15.9	15.0
80	28.8	25.8
70	52.6	46.6
60	68.8	60.2
50	188	159
40	467	367
20	3990	2560
10	$60.9.10^4$	$31.8.10^4$
0	$79.6.10^5$	$28.3.10^5$

Aluminum oxide (Al_2O_3) + Ferric oxide (Fe_2O_3)

Cirilli, 1947

%	χ 600°	850°	1000°
-	-0.38	-0.35	-0.29
-	9.0	8.2	8.7
20	22.0	15.1	19.0
33	33.2	27.1	21.2
50	160.3	113.5	24.6
60	307.5	285.9	-
67	608.3	395.0	28.2
70	380.1	311.3	-
75	60.0	37.8	30.0
86	21.8	24.3	48.0
100	20.7	21.3	-

Selwood, Lyon and Ellis, 1951

t	χ	t	χ	t	χ
		2.0 %			
27	- 3.3	- 89	- 5.2	-145	- 7.5
-43	- 4.4	-135	- 6.8	-187	- 9.5
		5.0 %			
27	- 6.1	- 73	- 9.1	-125	-12.7
-43	- 7.8	- 89	- 9.9	-145	-14.6
-58	- 8.2	-106	-11.0	-187	-19.7
		10.5 %			
27	-11.8	- 73	-16.7	-125	-21.1
-43	-15.1	- 89	-18.7	-145	-25.0
-58	-15.8	-106	-19.7	-187	-32.5
		20.0 %			
27	-15.8	- 73	-22.3	-125	-27.3
-43	-20.5	- 89	-22.9	-145	-31.0
-58	-21.8	-106	-25.1	-187	-38.9
		40.0 %			
27	-17.7	- 73	-21.5	-125	-25.6
-43	-20.0	- 89	-22.4	-145	-27.5
-58	-20.8	-106	-22.9	-187	-34.2
	80 %				
27	-37.1	-187	-44.2		

β form

Cirilli and Brisi, 1951

mol%	χ	mol%	χ
10	-17.6	55	-66.6
17	22.0	65	69.2
25	23.7	75	64.8
33	27.4	85	52.2
40	33.2	90	51.3
45	47.7	100	32.5
50	52.4		

Aluminum oxide (Al_2O_3) + Lanthanum oxide (La_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	%	f.t.
100	2315	40	1875
80	2040	20	1935
70	1900	0	2050
60	1875		

E : 1870°

Aluminum oxide (Al_2O_3) + Gadolinium oxide (Ga_2O_3)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	%	f.t.
100	1740	40	1885
80	1780	20	1950
60	1825	0	2050

Aluminum oxide (Al_2O_3) + Ferosoferric oxide (Fe_3O_4)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	%	f.t.
100	1580	40	1810
80	1675	20	1920
60	1760	0	2050

Aluminum oxide (Al_2O_3) + Mangane oxide (Mn_3O_4)

Wartenberg and Reusch, 1932 (fig.)

%	f.t.	%	f.t.
100	1560	40	1820
90	1590 E	30	1800
80	1630	20	1880
60	1760	10	2100
		0	2050

Yttrium oxide (Y_2O_3) + Uranium oxide (U_3O_8)

Hund, Peetz and Kottenhalm, 1955

a	b	a	b
50 mol%		25°	38.1 mol%
80	-0.08	81	0.18
82	-0.07	85	.21
85	-0.03	88	.24
89	+0.02	93	.28
95	.07	98	.32
98	.13	104	.38
103	.20	113	.48
108	.28	125	.59
115	.38	142	.74
122	.49	159	.91
129	.62	182	1.09
140	.81	195	.20
161	.18	200	.24
174	.42	203	.26
191	.71	215	.36
200	.84	224	.43
211	2.07	236	.54
	26.5 mol%		
80	0.78	145	2.07
85	.86	163	.44
90	.98	175	.63
97	1.10	184	.81
103	.22	202	3.06
112	.40	226	.36
122	.61	287	4.12
131	.80		

a = temperature coefficient = $1/T.10^5$
b = lg resistivity (ohm/cm)

mol%	lattice constant (in Å)	d
	25°	
72.7	5.376	-
60.8	.376	-
55.4	.368	-
50.0	.361	8.92
45.9	.353	-
42.9	.346	-
38.1	.338	8.31
34.3	.330	-
30.6	.322	-
26.5	.317	7.46
22.2	.320	-
18.1	.329	-
13.3	.329	-

Neodymium oxide (Nd_2O_3) + Uranium oxide (U_3O_8)

Hund and Peetz, 1953

mol%	lattice constant (in Å)
91.8	5.434
85.7	.434
78.8	.438
66.6	.445
57.1	.448
45.3	.469
40.0	.470

Praseodymium sesquioxide (Pr_2O_3) + Yttrium oxide (Y_2O_3)

Mc Cullough and Britton, 1952

mol%	lattice constant (in Å)
0	5.570
5.0	.561
10.0	.545
19.9	.527
35.0	.480
50.0	.437
65.0	.397
80.0	.358
87.5	.355
90.0	.329
100	.304

Lanthanumoxide (La_2O_3) + Uranium oxide (U_3O_8)

Hund and Peetz, 1953

mol%	lattice constant (in Å)
88.8	5.472
82.3	.473
76.5	.481
75.0	.485
63.9	.516
57.1	.531
48.4	.554
40	.586

Samarium oxide (Sm_2O_3) + Uranium oxide (U_3O_8)

Hund and Peetz, 1953

mol%	lattice constant (in Å)
80.0	5.415
66.6	.414
49.9	.411

Erbium oxide (Er_2O_3) + Uranium oxide (U_3O_8)

Hund and Peetz, 1952

mol%	d
	25°
64.2	10.40
50.5	10.00
33.3	9.58
25	8.87

Chromium sesquioxide (Cr_2O_3) + Ferric oxide (Fe_2O_3)

Cirilli, 1947

mol%	χ		
	600°	850°	1000°
0	27.2	25.7	25.7
30	55.0	28.2	26.6
50	62.2	33.1	25.3
65	56.2	25.6	22.2
75	56.0	24.0	23.7
85	48.3	23.7	21.8
100	20.7	21.3	-

Milligan and Merten, 1947

X - ray diffraction analysis.

Ferric oxide (Fe_2O_3) + Cobaltic oxide (Co_3O_4)

Robin and Benard, 1952 (fig.)

%	tr.t. 1	tr.t. 2
28.6	1000	-
33.4	500 - 820	-
48.3	500 - 780	-
51.2	800	1000
63.4	880	960
73.7	870	940
84.0	820	930 - 1000
90.0	500 - 790	920 - 940
100.0	-	900

Tantalum pentoxide (Ta_2O_5) + Niobium pentoxide (Nb_2O_5)

Schäfer, Dürkop and Jori, 1954

mol%	tr.t. mixed crystals
61	1200
67	1050
84	900

Ferrosoferric oxide (Fe_3O_4) + Mangane oxide (Mn_3O_4)

Sullivan and Mauer, 1950

%	tr.t. begins	tr.t. ends
65	385	510
70	405	605
80	575	820
90	950	1030
100	1070	1170

Uranium oxide (U_3O_8) + Praseodymium oxide (Pr_6O_{11})

Hund and Peetz, 1952

mol%	d
	25°
100	6.70
73.9	7.12
53.7	7.77
33.2	8.56
21.0	9.40
0	8.34

Sodium sulfide (Na_2S) + Cuprous sulfide (Cu_2S)

Friedrich, 1914

%	f.t.	E	%	f.t.	E
0	920	-	67	700	-
5	870	-	70	685	600
10	810	500	72	675	600
20	750	507	75	650	597
30	685	505	78	-	600
40	600	505	80	625	595
50	(530)	507	85	690	598
52	-	510	90	800	598
55	560	505	95	940	590
60	650	507	97.5	1040	595
65	680	508	100	1125	-

Sodium sulfide (Na_2S) + Ferrous sulfide (FeS)

Steck, Slavin and Ralston, 1929

%	f.t.	E
98.6	1174	703 - 688
93.5	1088	694 - 674
86.4	-	699 - 668
86.0	907	699 - 673
81.6	850	660 - 647
70.3	701	665 - 656
64.9	696	638 - 613
60.7	695	625 - 549
50.6	-	588 - 560
40.7	653	- -
0	970	

Thallium sulfide (Tl_2S) + Silver sulfide (Ag_2S)

Huber, 1921

%	f.t.	E	min.	tr.t.	min.
100	781	-	-	175	-
90	668	-	-	358	1
80	580	-	-	358	1.5
70	490	-	-	358	2
65	413	306	2	363	-
59.8	397	310	6.5	356	-
46.2	312.5	307.5	15	-	-
40	352	307	14	-	-
30	365	307	11	-	-
27.4	375	308	9.5	-	-
18.7	406	307.5	8	-	-
0	448	-		-	-

Thallium sulfide (Tl_2S) + Lead sulfide (PbS)

Canneri and Fernandes, 1925

mol%	f.t.	E	mol%	f.t.	E
0	448.5	-	45	463	282
2	440	-	60	596	282
10	417	282	70	894	282
20	381	282	80	981	-
30	340	282	90	1058	-
40	-	282	100	1108	-

Cuprous sulfide (Cu_2S) + Silver sulfide (Ag_2S)

Friedrich, 1907

%	f.t.	%	f.t.
100	835	40.5	791
91.4	749	30.4	876
80.9	698	20	945
70.5	677	10.2	1050
60.5	688	0	1121
50.3	719		

Urazov and Chelidze, 1941

%	f.t.	%	f.t.
100	835	40	814
90	714	30	900
80	667	20	966
70	650	10	1052
60	657	0	1150
50	619		

Cuprous sulfide (Cu_2S) + Cupric sulfide (CuS)

Buerger, 1941 (fig.)

%	tr.t.	%	tr.t.
0.5	300	9	90
1	120	13	78
1.5	110	16.5	82
5	105	20	100
6	20	24	140
7	140		

Ramdohr, 1943 (fig.)

%	tr.t.	%	tr.t.
0.5	300	10	90
1	150	13	75
2	103	15	80
5	103	20	115
6	20	22	140
8	140		

Cuprous sulfide (Cu_2S) + Lead sulfide (PbS)

Friedrich, 1907

%	f.t.	E	min.
100	1114	-	-
89.2	987	535	13
79.60	874	540	20
68.46	764	535	28
60	662	536	53
50.14	560	543	82
41.37	656	553	72
31.28	781	553	40
21.38	909	540	28
12.27	1033	535	11
0	1121	-	-

Cuprous sulfide (Cu_2S) + Zinc sulfide (ZnS)

Friedrich , 1912

%	f.t.	E	%	f.t.	E
0	1118	-	15	1183	1142
1	1136	1107	30	1226	-
2	1135	1114	40	1301	-
5	1148	1120	50	1370	-
10	1175	1149	60	1423	-
15	1194	1120	70	1501	1164
			80	1570	1154

Cuprous sulfide (Cu_2S) + Ferrous sulfide (FeS)

Schad and Bornemann, 1916

%	f.t.	E	min.	tr.t.	min.
100	1160	-	3.2	904	4.0
80.1	1052	973	3.9	908	4.1
75.1	1020	973	6.1	916	5.6
70.9	993	974	8.1	914	5.6
67.6	978	975	9.0	917	6.4
64.6	-	975	4.0	914	7.4
62.9	-	979	7.2	914	8.2
60.5	-	979	7.2	913	8.0
58.2	-	973	6.2	915	7.8
55.1	-	974	1.2	910	2.8
52.5	980	975	-	897	-
49.5	986	975	-	855	-
48.9	996	975	-	819	-
47.7	995	975	-	811	-
45.7	1008	975	-	796	-
45.6	1005	975	-	784	-
43.3	1015	975	-	747	-
40.9	1025	975	-	708	-
39.6	1030	975	-	690	-
36.4	1037	-	-	666	-
36.0	1054	-	-	630	-
33.5	1064	-	-	589	-
31.4	1066	-	-	532	-
29.3	1079	-	-	552	-
29.3	1082	-	-	522	-
27.2	1090	-	-	532	-
24.6	1099	-	-	-	-
25.2	1097	-	-	-	-
22.8	1102	-	-	-	-
21.4	1106	-	-	-	-
18.7	1104	-	-	-	-
15.2	1113	-	-	-	-
10.3	1130	-	-	-	-
0	1132	-	-	-	-

%	tr.t. I	tr.t. II	tr.t. III	tr.t. IV
80.1	220	140	-	-
75.1	-	138	-	-
70.1	-	140	-	-
67.6	230	-	-	55
64.6	236	140	88	-
62.9	-	140	70	-
60.5	-	-	88	-
58.2	217	130	84	-
55.1	-	140	-	60
52.5	-	"	-	60
49.5	-	"	-	42
48.9	226	"	-	50
47.7	-	135	85	-
45.7	-	127	95	35
45.6	224	132	86	-
43.3	-	138	90	-
40.9	-	139	75	-
39.6	-	130	90	-
38.4	-	134	90	-
36.0	237	140	85	-
33.5	238	134	94	-
31.4	250	140	98	-
29.3	-	150	94	-
29.3	-	138	90	-
27.2	204	152	88	-
25.6	210	140	80	-
25.2	-	-	-	-
22.8	-	139	85	-
21.4	133	81	-	-
18.7	-	135	75	-
15.2	-	-	78	-
10.3	-	-	110	60
0	264	104	108	58

Carpenter and Hayward, 1923

%	f.t.	E	tr.t.
100	1163	-	-
90	1140 - 1130	1005	948
85	1100	1000	952
80	1070 - 1076	1000	950
75	1045	995	950
72.5	1007	-	950
70	1000	-	952
69	1006	-	952
68	1005	-	950
67.5	995	-	950
65	996	-	950
62.5	995	-	-
60	998	-	951
50	1008	990	952
45	1042	-	mixed crystals
40	1060	-	"
35	1062	-	"
30	1070 - 70	-	"
20	1100	-	"
10	1115 - 10	-	"
0	1128	-	"

Srivalin and Essin, 1952

%	m.t.	f.t.	%	m.t.	f.t.
0	1120	1120	65	960	970
12.5	1090	1120	80	960	1040
25	1050	1080	87.5	1000	1090
42	960	1020	98	1130	1130
50	960	960	100	1170	1170

mol%	e (in millivolts)	mol%	e (in millivolts)
0	0	73	145
15	57	80	154
31	88	88	158
44	108	94	169
55	119	100	174
64	137		

Cuprous sulfide (Cu_2S) + Nickel subsulfide (Ni_2S)

Friedrich, 1914

%	f.t.	E	tr.t.
0	1125	-	-
1	1115	-	-
5	1085	575	525
10	1030	"	"
15	1020	"	"
20	1000	"	"
25	983	"	"
30	978	"	"
35	970	"	"
40	965	"	"
45	960	"	"
50	942	"	"
55	918	"	"
60	890	"	"
65	855	"	"
70	810	"	"
75	760	"	"
80	710	"	"
85	665	"	"
87.5	638	"	"
90	620	"	"
93.5	-	"	"
95	607	"	"
97.5	630	"	"
99	640	"	"
100	645	577	524

Cuprous sulfide (Cu_2S) + Nickel sulfide (Ni_3S_2)

Hayward, 1915

%	f.t.	E
0	1130	-
2.5	1086	-
5	1068	-
10	1043	727
20	1008	728
30	982	731
40	953	731
50	925	729
60	884	728-8
70	820	726
72.5	798	726
75	773	728
80	727-6	-
90	739	728
95	768	(740)
100	794	-

Friedrich, 1914

%	f.t.	E	tr.t.
0	1125	-	-
1	1113	545	(720)
2	1102	540	"
5	1080	543	"
10	1067	545	-
15	1010	543	722
20	985	545	"
25	971	547	725
30	952	543	728
35	947	542	"
40	938	543	727
45	923	543	725
50	915	546	"
55	900	545	"
60	880	545	727
65	852	543	728
70	824	543	"
75	804	544	"
77.5	798	542	"
80	784	543	726
82.5	775	545	725
85	753	545	"
87.5	(730)	545	727
90	-	545	725
91	-	547	728
92.5	750	547	"
95	760	544	"
97.5	772	544	"
99	785	545	727
100	790	-	723

Cuprous sulfide (Cu_2S) + Antimonium sulfide (Sb_2S_3)

Parravano and de Cesaris, 1912

%	f.t.	E	min.	
10.94	989	558	60	
16.48	923	572	120	
25.02	810	580	180	
33.58	714	590	255	
39.94	-	610	285	
41.38	610	-	300	
42.24	"	-	-	(1+1)
43.66	"	-	-	
44.29	"	-	-	
45.40	"	535	50	
47.37	602	535	60	
49.73	592	532	60	(3+1)
52.94	580	532	105	
58.10	564	530	180	
63.34	552	540	390	
66.67	548	542	420	
67.90	542	542	450	
72.00	540	450	45	
75	540	460	90	
79	532	468	120	
83.34	515	478	180	
89.0	-	490	255	
91.67	500	490	225	
96.67	528	474	75	

Silver sulfide (Ag_2S) + Lead sulfide (PbS)

Friedrich, 1907

%	f.t.	E	tr.t.
100	1114	-	-
95	1073	-	-
89.7	1027	-	-
80	960	-	-
75	928	626	170
70	903	635	-
60.2	841	"	180
50.3	781	"	170
45	738	615	-
40	707	635	-
30	658	625	173
16.6	667	625	172
15	691	-	-
10	739	-	-
0	835	-	-

Silver sulfide (Ag_2S) + Zinc sulfide (ZnS)

Friedrich, 1912

%	f.t.	E	%	f.t.	E
0	843	-	1.5	818	792
1.7	830	807	2.7	809	787
3.9	848	797	4	848	802
7	848	799	7	957	-
7.9	947	803	10	1041	803
30	1031	810	17	1192	808
			30	1300	-
			50	1413	792
			69.3	1508	796
			80.8	1574	805
			100	1670	-

Silver sulfide (AgS) + Ferrous sulfide (FeS)

Scholn, 1911

%	f.t.	E	tr.t.
0	812	-	175
2.5	758	560	173
5	714	611	172
7.5	680	626	183
10	629	616	172
15	678	627	173
10	734	611	172
30	827	617	176
40	911	615	177
50	985	615	179
60	1046	619	178
70	1088	613	177
80	1155	612	186
90	1170	615	175
100	1171	-	-

Silver sulfide (AgS) + Antimonium sulfide (Sb_2S_3)

Guinchant and Chrétien, 1904

%	f.t.
100	540
95.30	522
90.64	505
86.88	487
82.92	462

Jaeger and van Klooster, 1912

%	f.t.	E	min.	
100	546	-	-	
96.27	529	-	-	
92.44	510	443	10	
88.50	499	450	30	
84.45	479	447	40	
80.30	462	453	90	
77.57	-	451	120	
76.02	460	452	85	
71.61	482	447	30	
67.08	498	451	10	
62.41	506	-	-	(1+1)
57.59	509	-	-	
52.64	500	453	90	
47.17	479	454	120	
43.84	469	458	250	
42.77	-	458	270	
42.24	464	455	200	
36.78	475	454	35	
31.16	482	-	-	
25.35	470	463	380	(3+1)
24.02	-	464	395	
21.35	466	464	260	
19.34	505	464	245	
13.11	595	463	170	
10.18	647	458	100	
6.67	710	463	70	
0	842	-	-	

Calcium sulfide (CaS) + Ferrous sulfide (FeS)

Vogel and Heumann, 1941 (fig.)

%	f.t.	E
100	1200	-
90	1160	1120
80	1120	1120
70	1270	1120
60	1425	1120

Zinc sulfide (ZnS) + Lead sulfide (PbS)

Friedrich, 1912

%	f.t.	E	%	f.t.	E
100	1114	-	100	1120	-
98	1096	-	97.5	1083	1048
96.35	1069	-	87.5	1122	1052
95.9	-	1044	80	1273	-
92	-	1044	98	1080	-
90.7	1088	1041	84	1154	1052
88	1132	1051	70	1363	1042
84	-	1051	60	1408	1047
81.3	1277	1051			
81.0	1277	1046			
75.0	1322	1046			

Zinc sulfide (ZnS) + Manganese sulfide (MnS)

Kröger, 1939

linear with weight % - hexagonal a (in Å)

0 % 3.811 Å 100 % 3.976 Å

Juza, Rabenau and Pascher, 1956

mol%	lattice constant in (Å) a	c	molar vol/cc
	Blende structure		
0	5.404	23.91	-
5	.411	24.00	-
10	.419	24.11	-
	Wurtzite structure		
10	3.833	6.256	24.12
20	.849	.277	24.40
31	.866	.305	24.73
50	.895	.346	25.25
64	.885	.329	25.06
65	.883	.327	25.04
80	.884	.330	25.06
	NaCl- structure		
64	5.210	-	21.42
65	.209	-	21.41
80	.211	-	21.44
89	.209	-	21.41
100	.209	-	21.41

mol%	$\varkappa$	mol%	$\varkappa$
60	2.10^{-12}	90	9.10^{-8}
65	7.10^{-11}	96	2.10^{-7}
70	6.10^{-10}	98	3.10^{-7}
80	7.10^{-8}	100	4.10^{-7}

Zinc sulfide (ZnS) + Ferrous sulfide (FeS)

Friedrich, 1912

%	f.t.	E α	E β
100	1188	-	-
95.0	-	1180	1162
92.7	1222	1178	1161
90.5	1253	1175	1164
79.9	1418	1180	-
63.9	1511	1185	1160
47.7	1550	1180	1143
43.8	1568	-	1162

Stannous sulfide (SnS) + Lead sulfide (PbS)

Heike, 1912

%	f.t.	tr.t.	%	f.t.	tr.t.
100	1106	-	62.6	910	898
98	1100	-	50	886	-
95	1086	-	40	879	-
93	1075	-	30	891	-
90	1049	-	20	880	-
85	1017	-	10	880	-
80	1005	875	0	880	-
70	946	900			

Stannous sulfide (SnS) + Ferrous sulfide (FeS)

Haan, 1913

%	f.t.	m.t.	%	f.t.	m.t.
0	870	-	60	1024	770
5	831	785	70	1070	797
15	-	784	80	-	796
20	821	785	90	1150	796
25	846	785	95	-	786
40	929	794	100	1188	777
50	958	760			

Stannous sulfide (SnS) + Antimonium sulfide (Sb_2S_3)

Parravano and de Cesaris, 1912

%	f.t.	E	min.	tr.t.	min.
100	545	-	-	-	-
95	528	-	-	460	120
89.20	500	-	-	466	150
85	485	-	-	460	225
81.80	-	-	-	470	300
75	485	-	90	460	105
71.40	490	480	-	460	-
70.43	500	480	105	460	75
69.15	518	485	-	465	-
65.24	545	485	-	465	-
52.69	568	481	75	458	60 (1+1)
60.10	590	480	60	460	60
56.60	612	480	60	450	30
48.60	650	474	75	460	45
45.20	674	470	60	455	45
39.56	694	480	45	462	45
33.15	725	474	60	458	30
27.20	760	466	45	-	-
19.62	790	470	30	-	-
16.66	800	470	-	-	-
10.64	830	-	-	-	-
0	850	-	-	-	-

Wagenmann, 1912

%	f.t.	tr.t.	%	f.t.	tr.t.
100	540	-	39.75	680	576
89.90	480	430	24.81	944	579
81.86	440	422	15.10	1025	576
77.63	480	425	11.78	1044	574
69.91	530	428	7.89	1073	576
58.94	560	424	4.03	1084	-
49.02	580	535	0	1109	-

Jaeger and van Klooster, 1912

%	f.t.	m.t.	min.
100	546	-	-
96.4	535	454	10
92.7	519	464	30
89.9	509	481	90
84.9	496	496	220
80.9	504	480	80
76.7	518	480	90
73.8	527 (5+4)	470	70
70.6	538	464	60
67.9	550	467	40
63.2	564	450	-
58.5	570	490	-
56.5	574	555	-
54.5	582	560	-
53.0	590	565	-
50.9	595	564	-
48.4	605 (2+1)	573	-
46.3	630	610	-
44.6	665	606	-
41.3	690	600	-
37.7		597	-
36.1	- (3+2)	-	-
31.9		586	-
26.0	-	582	-
22.0	-	-	-
13.5	-	547	-
0	1120	-	-

%	tr.I	min.	tr.II	min.
58.5	430	30	-	-
56.5	519	-	-	-
54.5	510	-	400	20
53.0	522	-	391	10
50.9	504	-	392	-
48.4	522	-	-	-
46.3	565	-	-	-
41.3	540	-	470	-
37.7	-	-	-	-
36.1	-	-	-	-
31.9	-	-	-	-
26.0	-	-	-	-
22.0	-	-	-	-
13.5	-	-	-	-
0	-	-	-	-

Lead sulfide (PbS) + Ferrous sulfide (FeS)

Friedrich, 1907

%	f.t.	E	min.
0	1114	-	-
5.1	1063	866	3.8
7.7	1042	868	6.0
10.4	1015	867	7.5
20.6	929	869	14
32.8	885	864	23
42.1	955	863	16
52.0	1033	863	12
69.4	1105	854	8.0
72.4	1117	854	6.0
93.9	1170	852	1.8
100	1187	-	-

Lead sulfide (PbS) + Antimonium sulfide (Sb_2S_3)

Guinchant and Chrétien, 1904

%	f.t.
100	540
95.08	522
92.10	512
86.88	490
33.11	467

Pélabon, 1913

mol%	f.t.
100	550
78	482 E
50	568 tr.t.
40	610
0	1015

Iitsuka, 1921

%	f.t.	tr.t.I	min.
100	516.3	-	-
95	490.9	-	-
90	465.9	-	-
85	440.4	-	-
80	470.2	-	-
70	518.7	-	-
65	568.5	548	10
60	594.6	546.1	20
55	613.9	543.9	15
50	656.7	609.4	40
48	662.4	610.3	45
43	671.8	611.8	10
40	668.3	596.8	40
35	659.9	591.3	40
33	631.3	590.3	25
30	610.7	592.4	10
25	576.9	-	-
20	570.5	-	-
17	585.0	-	-
15	611.6	-	-
10	677.0	-	-
5	771.5	-	-
0	1051	-	-

%	f.t.	tr.t.II	min.	tr.t.III	min.
100	516.3	-	-	425.4	5
95	490.9	-	-	430.4	20
90	465.9	-	-	427.9	30
85	440.4	-	-	428.2	18
80	470.2	-	-	428.3	10
70	518.7	-	-	430.0	5
65	568.5	-	-	-	-
60	594.6	-	-	-	-
55	613.9	514.9	15	-	-
50	656.7	517.7	30	-	-
48	662.4	511.3	30	-	-
43	671.8	-	-	-	-
40	668.3	468.6	20	-	-
35	659.9	466.8	20	-	-
33	631.3	471.8	15	-	-
30	610.7	472.9	15	553.2	15
25	576.9	460.0	10	557.4	25
20	570.5	-	-	545.7	40
17	585.0	-	-	564.4	32
15	611.6	-	-	558.4	50
10	677.0	-	-	554.2	35
5	771.5	-	-	549.9	20
0	1051	-	-	552.2	10
				-	-
(1+1)	(3+2)	(2+1)	(5+2)		

Pyrrhotite (FeSx) + Pentlandite (FeNiS)

Newhouse, 1927

%	f.t.	m.t.	%	f.t.	m.t.
100	886	-	40	1144	$C_I + C_{II}$
90	1031	854	35	1161	at
80	1073	836	25	1170	890°
70	1084	890	15	1187	
60	1103	878	10	1201	
50	1090	865	0	1192	
45	1126	889			

Manganese sulfide (MnS) + Ferrous sulfide (FeS)

Shibata, 1928

%	f.t.	E	%	f.t.
100	1175	-	74.45	1293
96.79	1168	1164	69.99	1339
95.13	-	"	65.05	1467
93.97	-	"	60.45	1383
93.53	-	"	51.90	1428
90.40	1130	1165	41.13	1520
89.61	1190	1163	30.91	1570
83.06	1202	"	18.95	1602
82.94	1230	1162	9.63	1610
81.17	1250	"		

Silver selenide(Ag_2Se) + Antimonium selenide (Sb_2Se_3)

Pélabon, 1908

mol%	f.t.	mol%	f.t.
0	880	-	650 (3+4)
5	790	-	573
20	540	100	615

Silver selenide(Ag_2Se) + Bismuth selenide (Bi_2Se_3)

Pélabon, 1908

mol%	f.t.
0	880
5	790
-	670 E
-	770 (3+4)
100	718

Zinc selenide (ZnSe) + Manganese selenide (MnSe)

Juza, Rabenau and Pascher, 1956

mol%	lattice constant in (Å) a	c	molar vol/cc
	Blende structure		
0	5.650	-	27.33
5	.661	-	27.49
10	.671	-	27.63
20	.692	-	27.94
31	.713	-	28.25
35	.723	-	28.41
70	.760	-	28.96
80	.760	-	28.96
90	.760	-	28.96

	Wurtzite structure		
35	4.048	6.615	28.44
40	.056	.628	28.61
50	.072	.652	28.94
70	.072	.652	28.94
80	.072	.652	28.94
90	.072	.652	28.94
	NaCl structure		
50	5.448	-	24.50
70	.448	-	24.48
80	.447	-	24.48
90	.447	-	24.50
100	.445	-	24.46

mol%	$\varkappa$	mol%	$\varkappa$
50	1.10^{-12}	80	4.10^{-7}
60	8.10^{-10}	95	2.10^{-6}
70	7.10^{-8}	100	4.10^{-6}

Zinc telluride (ZnTe) + Manganese telluride (MnTe)

Juza, Rabenau and Pascher, 1956

mol%	lattice constant in (Å) a	c	molar vol/cc
	Blende structure		
0	6.089	-	34.20
10	.105	-	34.47
20	.125	-	34.81
30	.145	-	35.15
40	.162	-	35.45
50	.182	-	35.79
60	.198	-	36.07
70	.219	-	36.44
80	.238	-	36.77
90	.250	-	36.99
95	.249	-	36.97
98	.249	-	36.97
90	.129	6.618	29.96
95	.128	.699	29.95
98	.127	.698	29.94
100	.127	.696	29.93

mol%	$\varkappa$	mol%	$\varkappa$
0	5.10^{-10}	60	6.10^{-10}
0.5	8.10^{-9}	70	1.10^{-10}
10	1.10^{-10}	80	9.10^{-9}
20	9.10^{-10}	90	1.10^{-5}
30	2.10^{-11}	95	4.10^{-5}
40	4.10^{-11}	98	7.10^{-5}
50	2.10^{-11}	100	8.10^{-5}

Nickel arsenide (NiAs) + Cobalt arsenide (CoAs)

Friedrich, 1912

%	f.t.	m.t.	tr.t.
0	970	945	481
1	973	945	-
2	975	945	-
5	981	945	497
10	992	952	518
20	1021	961	538
30	1045	970	590
40	1070	980	620
50	1093	1013	660
70	1128	1058	688
90	1164	1114	712
100	1181	1140	720

Tantalum carbide (TaC) + Titanium carbide (TiC)

Kovalski and Umanski, 1946

mol%	lattice constant a
0	4.4452
25	.4468
50	.3760
75	.3487
100	.3190

Tantalum carbide (TaC) + Zirconium carbide (ZrC)

Agte and Altertham, 1930

mol%	f.t.
100	3532
50	3772
33.3	3807
20	3932
11.1	3907
0	3872

Kovalski and Umanski, 1946

mol%	lattice constant a
0	4.4452
25	.4949
50	.5440
75	.6009
100	.6792

Tantalum carbide (TaC) + Hafnium carbide (HfC)

Agte and Altertham, 1930

mol%	f.t.
100	3887
25	3932
20	3942
16.5	3922
11	3912
0	3917

Tantalum carbide (TaC) + Tungsten carbide (WC)

Agte and Altertham, 1930

mol%	f.t.
100	2857
91	2777
87	2877
66.7	2947
50	3212
33.3	3407
20	3583
11.1	3702
0	3877

Tantalum carbide (TaC) + Niobium carbide (NbC)

Agte and Altertham, 1930

mol%	f.t.
100	3497
50	3627
20	3752
11	3827
0	3872

Zirconium carbide (ZrC) + Niobium carbide (NbC)

Agte and Altertham, 1930

mol%	f.t.
100	3532
50	3517
0	3497

Kovalski and Umanski, 1946

mol%	lattice constant a (in Å)
0	4.4600
25	.5110
50	.544
75	.6105
100	.6792

Zirconium carbide (ZrC) + Tungsten carbide (WC)

Agte and Alterthum, 1930

mol%	f.t.
100	2857
50	2857
20	2857
0	3532

Niobium carbide (NbC) + Tungsten carbide (WC)

Agte and Alterthum, 1930

mol%	f.t.
100	2857
50	3027
66.7	3167
0	3497

Niobium carbide (NbC) + Titanium carbide (TiC)

Kovalski and Umanski, 1946

mol%	lattice constant a (in Å)
0	4.4600
20	.4325
35	.4103
50	.3933
65	.3607
90	.3270
100	.3190

XLII. TWO OXY SALTS

Lithium hydroxide (LiOH) + Sodium hydroxide (NaOH)

Diogenov, 1953 (fig.)

mol%	f.t.	mol%	f.t.
0	477	61.5	238
25	400	70	220
50	300	94.5	299
57	254	100	320

Sodium hydroxide (NaOH) + Potassium hydroxide (KOH)

Neumann and Bergve, 1914

%	f.t.	%	f.t.
0	296	54.8	168.5
9.67	280	58.4	167.0
17.8	264	70.8	190.5
35.4	226	80	219.0
47	197	100	345.0

von Hevesy, 1910

wt%	mol%	f.t.	m.t.	tr.t. beginn.	tr.t. end
100	100	360	360	-	-
96.2	95.2	350	347	246	246
92.7	90.1	345	341	243	243
86.9	82.6	325	312	231	227
80.0	74.1	297	286	219	215
71.7	64.6	264	250	208	207
64.6	56.7	233	242	197	197
55.6	47.3	200	198	190	187
52.1	43.7	191	191	185	185
50.4	42.0	188	188	180	180
46.9	38.7	185	185	180	180
42.3	34.1	199	195		
37.1	29.7	213	203		
31.0	24.2	228	215		
23.9	18.4	245	232		
18.8	14.2	259	245		
6.7	4.8	292	281		
0	0	318.0	318		

Sodium hydroxide (NaOH) +Rubidium hydroxide (RbOH)

von Hevesy, 1910

wt%	f.t.	m.t.	tr.t. beginn.	tr.t. end	E	min.
100	301	301	245	245	-	-
95.0	280	276	200	189	-	-
91.4	270	241	172	172	241	30
85.9	255	241	171	171	"	119
83.1	242	241	171	171	"	157
75.2	267	241	171	171	"	83
68.0	272	241	170	170	"	20
60.6	278	240	171	182	240	32
53.2	268	266	188	200		
50.1	244	239	202	202		
48.2	237	235	207	204		
46.4	249	246	211	223		
38.8	261	254	227	236		
30.7	273	264	240	251		
22.4	284	280	254	266		
13.3	290	288	268	272		
6.8	305	300	276	272		
0	318	318	300	300		

Potassium hydroxide (KOH) + Rubidium hydroxide (RbOH)

von Hevesy, 1910

%	mol%	f.t.	m.t.	tr.t. beginn.	tr.t. end
0	0	360	360	248	248
4.8	3.0	382	380	247	247
12.6	4.0	395	393	248	246
21.0	12.8	399	399	244	213
26.2	16.2	398	397	243	243
35.9	23.7	375	371	236	235
44.1	31.0	327	324	230	228
48.8	34.4	306	304	216	215
51.7	37.1	309	307	215	215
60	48.1	321	317	225	222
68.6	54.5	334	327	233	231
77.3	44.0	344	339	240	239
79.9	66.8	348	344	242	242
86.4	77.7	351	350	248	248
90.9	85.5	350	350	248	247
95.5	92.1	332	328	247	245
100	100	301	301	244	244

Sodium nitrite ($NaNO_2$) + Potassium nitrite (KNO_2)

Amadori, 1913

mol%	f.t.	tr.t.	mol%	f.t.	tr.t.
0	282	160	50.00	248	-
3.53	270	135	54.88	262	-
8.27	264	110	65.43	300	-
16.87	246	-	76.43	340	-
25.80	225	-	87.96	383	-
35.10	219	-	100	440	(387)
44.78	233	-			

Ettinger, 1932

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	387	387	40	300.5	241
90	386	379	30	230.5	230.5
80	384.5	364.5	20	260.5	245.7
70	379.7	335.5	10	281.5	268.5
60	371	300	0	284	284.0
50	357	266.5			

Lithium metaborate ($LiBO_2$) + Sodium metaborate ($NaBO_2$)

van Klooster, 1910

%	f.t.	E	min.
0	843	-	-
10	808	651	90
20	765	650	130
30	734	654	180
40	-	655	300
50	-	650	400
60	726	647	250
70	795	648	150
80	874	646	50
90	919	648	50
100	966	-	-

Lithium metaborate ($LiBO_2$) + Potassium metaborate (KBO_2)

Bergman, Kislova and Posipaiko, 1955

mol%	f.t.	mol%	f.t.
100	947	40.0	624
95.0	920	35.0	660
90.0	900	30.0	692
85.0	884	20.0	764
75.0	846	15.0	790
65.0	790	12.5	794
55.0	718	10.0	798
50.0	672	7.5	810
45.0	590	5.0	820
42.5	604	0.0	843

E : 56 mol % 582° tr.t. 89 mol % 794°

Sodium metaborate ($NaBO_2$) + Potassium metaborate (KBO_2)

van Klooster, 1910

%	f.t.	m.t.	sat.t.
0	966	966	-
10	944	936	-
20	914	883	-
30	882	-	-
40	861	-	529
50	850	850	553
60	863	849	522
70	886	857	-
80	910	878	-
90	929	918	-
100	947	947	-

Magnesium aluminate ($MgAl_2O_4$) + Zinc aluminate ($ZnAl_2O_4$)

Kordes and Becker, 1949

mol%	d	n_D
	20°	
0.0	3.546	1.714
24.8	.810	.734
45.7	4.022	.751
71.2	.274	.772
84.1	.404	.782
92.3	.488	.787
100.0	.602	.795

Calcium aluminate ($Ca_3Al_2O_6$) + Strontium Aluminate ($Sr_3Al_2O_6$)

Carlson, 1955 (fig.)

mol%	lattice constant, Å
0	15.26
25	15.40
50	15.55
75	15.70
100	15.85

Magnesium ferrite ($MgFe_2O_4$) + Nickel ferrite ($NiFe_2O_4$)

Forestier and Vetter, 1939 (fig.)

mol %	t Curie point	mol %	t Curie point
100	590	40	445
90	555	30	410
80	540	20	380
70	515	10	350
60	485	0	320
50	475		

Zinc ferrite ($ZnFe_2O_4$) + Cadmium ferrite ($CdFe_2O_4$)

Chalyi and Rozhenko, 1956

%	Lattice constant a	%	Lattice constant a
0	8.426	70.51	8.561
19.29	8.459	78.20	8.595
23.00	8.466	82.70	8.466
28.49	8.481	85.66	8.617
37.40	8.491	100	8.680
54.44	8.527		

Zinc ferrite ($ZnFe_2O_4$) + Nickel ferrite ($NiFe_2O_4$)

Forestier and Vetter, 1939 (fig.)

mol %	t Curie point	mol %	t Curie point
100	590	40	200
90	525	30	1 125
80	475	20	50
70	410	10	-
60	350	0	-
50	270		

Cupric ferrite ($CuFe_2O_4$) + Nickel ferrite ($NiFe_2O_4$)

Forestier and Vetter, 1939 (fig.)

mol %	t Curie point	mol %	t Curie point
100	590	40	500
90	575	30	485
80	560	20	475
70	550	10	465
60	530	0	455
50	525		

Zinc ferrite ($ZnFe_2O_4$) + Cobalt ferrite ($CoFe_2O_4$)

Chalyi and Rozhenko, 1956

%	Lattice constant a	%	Lattice constant a
0	8.427	66.07	8.395
16.30	8.423	74.50	8.388
24.50	8.422	79.57	8.385
32.73	8.420	82.96	8.383
49.34	8.406	100	8.372

Sodium chlorate ($NaClO_3$) + Silver chlorate ($AgClO_3$)

Foote, 1902

t	mol % tetragonal		mol % cubic
12	49.96		14.33
25	51.81	C_1+C_2	18.45
35	52.08		21.73
50	52.90		26.56

Perucca, 1917

%	d	%	d
100	4.40	12.15	4.18
15.1	4.14	5.18	4.14
13.87	4.21	0	4.21

%	$(\alpha)_D$	%	$(\alpha)_D$
16.7	4.10	5.61	3.44
14.3	3.95	4.2	3.345
13.1	3.88	3.91	3.33
12.2	3.81	3.23	3.30
9.9	3.70	2.6	3.26
8.55	3.60	1.55	3.20
7.65	3.53	0	3.118
6.90	3.515		

Lithium nitrate ($LiNO_3$) + Sodium nitrate ($NaNO_3$)

Carveth, 1898

%	f.t.	%	f.t.
100.00	308	40	214
90	283	30	223
80	259	20	234
70	236	10	244
60	217	0	253
50	206		

Benrath and Tesche, 1920

t	ϰ (mhos)	t	ϰ (mhos)
	0%		
215	2.4	240	17.1
220	3.4	245	33.3
225	4.9	250	76.6
230	7.0	252	450
235	10.9		

	ϰ (mhos)				
t	1%	10%	20%	30%	40%
180	4.2	2.4	0.5	0.4	0.4
185	8.1	6.4	4.1	2.0	0.5
190	19.2	16.1	13.0	7.6	4.3
195	91.2	41.7	25.1	12.5	7.1
200	301.3	140.3	-	40.1	26.0
205	2500	4300	2377	1694	3596
t	50%	60%	70%	80%	90%
180	0.4	0.4	0.9	2.1	3.6
185	0.8	1.2	2.2	5.0	7.7
190	4.5	4.5	6.4	10.9	14.6
195	9.0	11.2	14.6	28.2	50.5
200	75.2	-	107	155	962
205	2317	-	1567	2239	5082
t	94%	95%	96%	97%	99%
180	-	-	-	-	5.1
185	-	-	-	1.0	11.0
190	-	-	-	2.0	12.0
195	-	-	0.6	3.6	63.6
200	-	-	1.7	15.1	1715
205	-	-	2.8	446	-
210	-	0.3	4.5	2100	-
215	0.4	1.4	8.5	-	-
220	0.9	7.6	20.1	-	-
225	-	26.2	57.1	-	-
230	1.8	61.5	195.9	-	-
235	-	112.7	870	-	-
240	4.5	184.1	-	-	-
245	-	470	-	-	-
250	8.1	-	-	-	-
260	32.7	-	-	-	-
265	93.8	-	-	-	-
270	939.7	-	-	-	-

t	ϰ	t	ϰ
	100%		
250	0.5	290	11.8
260	1.4	295	19.2
265	-	305	28.8
270	3.2	307	38.4
280	6.1	308	317

Blidin, 1941

mol%	f.t.	mol%	f.t.
8.1	250	43.2	200
15.9	243	46.7	210
20.6	236	49.5	216
24.9	230	54.4	224
28.7	222	56.7	230
32.0	216	62.9	245
35.2	210	69.6	260
38.2	202	75.8	272
40.8	196	89.0	295

E : 41.5 mol% 195°

Lehrman and Breslow, 1938

%	f.t.	E	%	f.t.	E
100.0	305.5	-	50.4	-	191.2
84.8	273	191.0	49.8	197	-
74.8	250	-	45.1	204	-
65.5	225	-	40.2	-	191.0
57.4	208	190.5	33.6	219	-
55.3	206	-	20.7	231	193.1
52.8	200	-	0.0	251.4	-

Protsenko, 1952 (fig.)

mol%	f.t.
0	252
25	220
43.75	196 E
75	260
100	306

Diogenov 1956.

mol%	f.t.	mol%	f.t.
100	308	46.5	199
98.7	307	45	196°E
95.8	300	42.5	200
91	290	40	203
88	284	36.5	208
84.5	277	32.5	214
81	270	28.5	220
77	262	25	225
72.5	252	20	231
69.5	246	15	237
65.5	238	10.5	243
62	231	7	247
57.7	221	4.5	250
53	212	1.5	252
50	206	0	254

Lithium nitrate ($LiNO_3$) + Potassium nitrate (KNO_3)

Carveth, 1898

%	f.t.	%	f.t.
0	253	60	143
11.11	242	65	132
25.00	224	70	139
33.33	212	75	186
40.00	200	80	223
45	188	85	255
50	174	90	290
55	161	100	337

Harkins and Clark, 1915

%	f.t.	%	f.t.
100	346.1	40	207.0
90	290.9	30	219.8
80	237.1	20	234.2
70	139.8	10	242.3
60	145.2	0	254.1
50	177.4		

Protsenko, 1952 (fig.)

mol%	f.t.	mol%	f.t.
0	252	58	120
25	215	75	250
50	140	100	337

Lithium nitrate ($LiNO_3$) + Rubidium nitrate ($RbNO_3$)

Pushin and Radoicic, 1937

mol%	f.t.	mol%	f.t.
0	255	60	177
10	245	68	154
20	225	70	167
30	195	80	217
35	179.5 E	90	261
40	185	100	311
50	191 (1+1)		

Protsenko and Dionisiev, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	252	68 E	155
20	220	80	225
35 E	178	100	310
50	193		

Lithium nitrate ($LiNO_3$) + Cesium nitrate ($CsNO_3$)

Pushin and Radoicic, 1937

mol%	f.t.	mol%	f.t.
100	418	43	174
90	375	40	187
80	320	30	212
70	264	20	232
60	220	10	244
50	190	0	255

Lithium nitrate ($LiNO_3$) + Thallium nitrate ($TlNO_3$)

Protsenko and Chelomov, 1953 (fig.)

mol%	f.t.	mol%	f.t.
100	206	50	175
75	150	25	225
70.5	132	0	255

Briscoe, Evans and Robinson, 1932

mol%	f.t.	m.t.	E	tr.t.
0	252	-	-	-
10.52	239	120	-	57
19.65	228	130	-	-
29.70	212	133.5	-	60
39.51	191	133	-	-
48.30	174	135	-	-
56.60	159	136	-	-
58.30	148	134	-	-
61.38	143	135	-	61
63.78 E	-	136 E	-	61
66.21	136.5	134	-	-
68.77	141	134.5	-	60
75.00	154	145	135	-
79.35	162	144	133	60
88.65	184	144	133.5	60.5
99.65	206	143.5	-	60.5

Lithium nitrate ($LiNO_3$) • Silver nitrate ($AgNO_3$)

Protsenko and Dionisiev, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	252	60	195
20	240	74 E	158
30	230	80	165
40	220	100	209

Palkin, 1926

mol%	f.t.	E	mol%	f.t.	E
100	208	-	53	202	172
95	199	172	33.5	224	171.5
90	191	172	20	237	172
85	184	172	16	239	172
80	178.5	171.5	10	244	172
75	171.5	172	5	243.7	172
65	186	171	0	249	-
60	193	170.5			

Cowen and Axon, 1956

mol %	$\varkappa$ 200°	250°	300°
100	6000	8300	10500
80	6000	8100	10600
70	5900	8100	10500
60	5900	8100	10600
40	5700	8000	10600
20	5400	7900	10700
10	5500	7900	10700
0	4900	7800	10600

Lithium nitrate ($LiNO_3$) + Strontium nitrate (SrN_2O_6)

Gromakov and Gromakova, 1953

mol%	f.t.	m.t.	
0	242	-	A
3	234	224	"
5	230	-	(1+1)
8	308	-	B
10	358	227	"
15	400	228	"
20	430	227	"
30	468	227	"
40	495	354	"
50	515	370	"
60	540	426	"
100	645	645	"

Tokareva and Bergman, 1956.

mol%	f.t.	mol%	f.t.
0	256	24.5	367
3.9	258	27.5	379
7.6	249 E	30.5	393
11.3	278	33.3	405
14.8	306	36.0	412
18.1	328	38.7	418
21.4	351		

Lithium nitrate ($LiNO_3$) + Barium nitrate (BaN_2O_6)

Harkins and Clark, 1915

%	f.t.	%	f.t.
0	254.1	44.80	403.58
17.37	303.17	55.80	436.6
32.13	364.4	100	595.53

Lithium nitrate ($LiNO_3$) + Cadmium nitrate (CdN_2O_6)

Protsenko, 1952 (fig.)

mol%	f.t.	mol%	f.t.
100	350	14.3	250
60	320	0	252
26.3	190 E		

Popovskaya and Protsenko, 1954

mol%	$\varkappa$ 200°	220°	240°	260°	280°	300°	320°
46.1	-	-	-	-	-	2060	2780
51.8	-	-	-	-	-	2420	3110
54.5	-	-	-	1520	2000	2620	3340
57.1	-	-	-	1640	2190	2860	3600
59.6	-	-	1290	1800	2400	3080	3810
62.0	-	920	1380	1900	2520	3190	3920
63.0	660	1080	1570	2140	2780	3450	4180
63.5	710	1140	1650	2230	2880	3570	4300
64.8	820	1290	1840	2410	3060	3720	4470
68.8	1010	1510	2080	2680	3340	4050	4830
70.9	-	1650	2220	2870	3580	4340	5110
73.0	-	1850	2460	3140	3900	4650	5460
75.0	-	2000	2640	3360	4900	4900	5690
78.8	-	-	2850	3610	4400	5170	6000
82.3	-	-	3180	3940	4760	5620	6500
85.7	-	-	3700	4520	5350	6200	7060
89.0	-	-	-	4700	5750	6800	7650
94.7	-	-	-	5670	6900	8100	9250
100.0	-	-	-	7950	9300	10700	12100

Sodium nitrate ($NaNO_3$) + Potassium nitrate (KNO_3)

Heterogeneous equilibria .

Schaffgotsch, 1857

%	f.t.	%	f.t.
0	313.1	50	236.3
10	297.7	54.3	228.8
20	281	60	225.6
37.26	262	70	229.8
37.29	248	70.4	250.3
40	248	80	280.4
41.1	243.6	90	311
43.9	241.6	100	338.3

Guthrie, 1884

%	f.t.
0	305
67.10	215 E

Carveth, 1898

%	f.t.	%	f.t.
0	308	55	221
10	293	60	228
20	276	70	248
30	259	80	277
40	240	90	308
50	224	100	337

Hissink, 1899

mol%			
L	C	mol%	tr.t.
30.1	25	100	127
60.2	89.2	98.43	126.4-127.4
40.3	20.6	96.5	121-124
62.5	87.6		

Fawsitt, 1908

%	f.t.	%	f.t.
0	308	65	237
20	278	76	263
40	240	100	339
54.5	224		

Goodwin and Mailey, 1908

mol%	f.t.
0	305.6
20	274
50	220
80	284.0
100	333.7

Menzies and Dutt, 1911

mol%	f.t.	mol%	f.t.
0	315.1	60	242.9
10	290.1	70	262.5
20	274.5	80	289.3
30	255.0	90	316.7
40	237.6	100	346.3
50	228.1		

Amadori, 1912-13

%	f.t.	tr.t.	
0	310	-	
11.67	287	-	
22.92	270	-	
28.39	260	-	
33.76	250	-	
44.22	228	-	
49.32	219	-	
54.32	219	85	mixed crystals gap
59.25	223	"	"
64.08	230	"	"
73.51	254	"	"
78.10	267	"	"
82.63	285	"	"
91.72	310	"	"
95.76	323	102	
97.89	330	116	
100	336	124	

Harkins and Clark, 1915

%	f.t.	%	f.t.
0	314.1	64.09	233.8
11.68	290.9	73.98	257.5
23	274.1	82.62	287.3
33.78	253.8	91.03	313.7
44.21	235.2	100	346.3
54.31	222.4		

Quartaroli, 1920

%	f.t.	%	f.t.
0	306	55	220 E
5	297	60	225
10	287	65	231
15	280	70	243
20	273	75	255
25	265	80	270
30	254	85	283
35	245	90	297
40	237	95	312
45	231	100	336
50	225		

Madgin and Briscoe, 1923

%			
L	C	L	C
19.1	15.0	(50.1)	(48.5)
20.1	17.3	54.9	55.2
22.5	18.4	(57.75)	(59.25)
26.1	22.3	59.9	62.9
27.5	29.5	65.25	69.0
32.5	31.0	65.5	70.0
34.5	35.2	70.4	75.1
38.5	37.7	75.0	79.8
40.6	42.3	81.0	85.5
44.4	42.3	85.6	90.7
46.8	43.0	90.4	93.4

Briscoe and Madgin, 1923

%	f.t.	%	f.t.
49.4	228.1	31.2	259.0
50.1	227.5	38.0	246.2
51.2	226.5	43.7	235.8
52.7	225.9	48.6	228.8
53.5	225.6	52.2	226.1
54.8	225.4	52.9	226.0
55.7	225.7	53.7	225.9
56.1	226.2	54.7	225.5
59.1	228.8	55.3	225.7
65.0	238.1	57.75	227.2
68.5	245.0	60.0	229.1
77.6	268.9	63.0	233.7
90.8	307.9	75.0	262.0
100.0	332.6	85.0	291.8
-	-	100	333.1

Tammann and Ruppelt, 1931

mol%	f.t.	m.t.	
		1	2
0	311	-	-
30	251	223	225
40	230	225	225
55	225	225	223
65	243	222	224
70	253	225	225
85	295	227	227
90	312	237	237
100	337	-	-

Laybourn and Madgin, 1934

%	f.t.	%	f.t.
50	227	80	277
60	229	90	307
70	250	100	340

Protsenko and Bergman, 1950

%	f.t.	%	f.t.
100	337	50	222
97.5	328	45	225
95	324	40	229
92.5	316	35	237
90	310	30	245
87.5	306	25	257
85	297	22.5	262
82.5	290	20	266
80	282	17.5	273
75	270	15	276
70	257	12.5	283
65	243	10	286
60	231	5	300
55	226	0	308

Kofler, 1955

%	f.t.	m.t.	tr.t.
0	308	-	-
10	294	258	195
20	278	242	205
30	260	231	205
40	240	226	195
50	226	223	175
55	222	222	165
60	228.5	223	-
70	249.5	229	-
80	278	238	-
90	305	265	-
100	335	-	127

Properties of phases

Goodwin and Mailey, 1907

t	d
0%	
342.6	1.893
400.8	.852
433.6	.828
452.4	.814
491.5	.789

$d_t = 1.835 (1 - 0.000379(t - 400))$

t	d
20 mol%	
286.2	1.927
332.4	.894
363.4	.870
412.0	.835
440.8	.816

$d_t = 1.844 (1 - 0.000400(t -400))$

t	d
50 mol%	
320.8	1.891
365.0	.858
414.5	.821
437.0	.804

$d_t = 1.832 (1 - 0.000408(t - 400))$

t	d
80 mol%	
312.6	1.894
383.6	.840
445.8	.793

$d_t = 1.828 (1 - 0.000407(t -400))$

t	d	t	d
100 mol%			
352.0	1.858	451.0	1.785
365.0	.853	468.0	.775
372.3	.847	497.0	.756
391.5	.832	504.0	.749
404.6	.823	535.0	.726
421.4	.809	553.9	.712
427.8	.804	554.0	.714
443.9	.794	564.0	.706

$d_t = 1.826 (1 - 0.000405(t - 400))$

Fawsitt, 1908

t	d	t	d
0%		20%	
324	1.903	278	1.930
333	.899	328	.894
353	.887	344	.884
393	.860	373	.864
40%		65%	
262	1.938	245	1.943
276	.928	253	.936
316	.897	289	.907
332	.883	340	.870
54.5%		76%	
229	1.963	276	1.920
236	.957	326	.880
269	.934	335	.873
286	.920		
291	.916		
309	.902		
331	.884		
346	.873		

Sandonnini, 1920

%	d	
	350°	400°
100	1.864	1.824
92.62	.865	.826
54.84	.870	.834
28.90	.875	.840
0	.884	.849

Guthrie, 1884

%	d
0	2.2028
67.10	.1328
100	.0469

Laybourn and Madgin, 1934

d	t	d	t	d	t
100%		90%		80%	
1.860	347	1.887	314	1.907	291
.855	354	.868	340	.877	332
.844	369	.852	362	.850	369
.831	387	.829	394	.825	404
.806	420	.802	431	.797	442
.792	440				
.776	462				
70%		60%		50%	
1.934	258	1.948	242	1.943	252
.914	286	.920	281	.905	305
.890	318	.892	320	.883	335
.859	360	.882	333	.838	398
.825	407	.842	388	.797	455
.797	446	.796	452		

Fawsitt, 1908

t	η	t	η
0%		20%	
324	2710	278	3250
333	2560	328	2480
353	2350	344	2340
393	2000	373	2170
40%		65%	
262	3680	245	4230
276	3240	253	4050
316	2470	289	3060
332	2370	340	2460
54.5%		76%	
229	4700	276	3380
236	4460	326	2590
269	3340	335	2420
286	2960		
291	2870		
309	2660		
331	2470		
346	2380		

%	t	η	t	η
0	328	2620	308	2960
20	298	2870	314	2660
40	260	3700	320	2440
54.5	244	4100	324	2520
65	257	3870	327	2550
76	283	3220	331	2500
100	359	2500	339	2830

Goodwin and Mailey, 1908

t	η	t	η
0%		50 mol%	
337.0	2538	328.0	2597
353.0	2293	335.0	2500
356.0	2283	337.3	2487
357.0	2267	370.0	2092
385.2	1992	382.0	1949
406.0	1779	468.0	1418
495.0	1321		
80 mol%		100 mol%	
345.0	2570	347.0	2793
347.0	2500	371.0	2352
358.5	2350	377.0	2298
369.5	2247	396.0	2141
388.0	2066	418.0	1890
441.0	1636	462.0	1589
		506.0	1418

Boardman, Palmer and Henman, 1955 (fig.)

mol%	σ 350°	σ 400°	mol%	σ 350°	σ 400°
0	96	82	60	80	66
20	93	78	80	78	64
40	82	70	100	77	63

Laybourn and Madgin, 1932

%	t.b.s.	%	t.b.s.
10	0.60	55	0.26
15	.77	60	.29
20	.88	65	.36
25	.93	70	.49
30	.91	75	.48
35	.72	80	.46
40	.49	85	.42
45	.40	90	.34
50	.32		

%	c.s. continuous loading a	b	c	intermittent loading a
0	22.5	-	24.5	17
10	133	116	148	119
25	259	182	-	217
40	154	115.5	346	140
55	31.5	11	115.5	20.5
70	147	63	259	126
85	126	52.5	245	91
100	53	-	77	41

t.b.s. = tranverse breaking strength (in Kg/cm_2)

c.s. = compression strength (in Kg/cm_2)

a = specimen annealed for 4 days

b = " " " 14 "

c = " unannealed

Bloom and Rhodes, 1956.

$n_D = a - bt$

mol%	t	a	b.10$_4$
80	300-400	1.472	1.7
60	280-400	1.465	1.5
40	300-400	1.462	1.4
20	300-400	1.459	1.2

Voskresenskaya, Jankovskaya and Anosov, 1948

54.3% m.t. = 224°

t	U C	U L
110.0	287	-
150.5	280	-
180.6	294	-
210.2	296	-
230.5	429	395
270.5	418	365
310.5	406	349
350.5	391	328
501.5	353	296

Benrath, 1908

t	0%	5%	κ.10⁶ 10%	20%	25%
180	-	0.4	1.2	2.3	2.6
185	-	-	-	-	3.0
190	-	0.7	1.5	3.3	3.7
195	-	-	-	-	4.3
200	-	1.3	2.1	5.5	5.6
205	-	-	-	-	8.1
210	0.4	3.3	6.4	8.9	11.3
215	-	-	-	13.0	12.7
216	-	-	-	-	-
217	-	-	-	-	-
218	-	-	-	-	-
219	-	-	-	-	18.3
220	0.7	5.1	10.7	17.2	22.2
221	-	-	-	-	-
222	-	-	-	-	70.5
223	-	-	-	-	220
225	-	-	-	26.0	-
227	-	-	-	-	-
228	-	-	-	39.8	-
229	-	-	-	56.5	-
230	1.1	7.6	17.5	134	-
231	-	-	-	450	-
235	-	-	21.2	-	-
240	1.5	10.3	29.5	-	-
245	-	-	51.3	-	-
247	-	-	85	-	-
248	-	-	118	-	-
249	-	-	177	-	-
250	2.4	13.1	304	-	-
251	-	-	-	-	-
255	-	-	-	-	-
260	3.6	-	-	-	-
265	-	-	-	-	-
270	5.9	-	-	-	-
271	-	-	-	-	-
272	-	-	-	-	-
273	-	-	-	-	-
280	7.8	-	-	-	-
285	-	-	-	-	-
290	14.0	-	-	-	-
292	-	-	-	-	-
294	-	-	-	-	-
295	18.8	-	-	-	-
300	24.8	-	-	-	-
304	50.0	-	-	-	-
305	60.2	-	-	-	-
306	76.5	-	-	-	-
307	142	-	-	-	-
308	200	-	-	-	-
t	**40%**	**50%**	**70%**	**80%**	**85%**
180	2.7	2.7	2.4	2.2	1.3
185	3.7	3.7	2.8	-	-
190	4.6	4.3	4.0	3.0	2.0
195	5.5	4.9	4.8	-	-
200	6.5	6.1	6.2	3.7	2.9
205	9.0	8.6	9.5	-	3.8
215	14.5	13.9	13.4	6.7	5.0
216	27.2	31.8	20.4	8.9	6.7
217	34.0	41.7	25.7	12.0	-
218	42.2	59.0	33.2	30.7	31.1
219	52.8	76.9	40.6	51.2	47.7
220	97	165	54.3	77.2	89.6
221	372	325	165	184	145
221	-	-	-	-	256
255	-	15.1	-	-	-
260	-	19.2	-	-	-
265	-	28.5	-	-	-
270	-	56.0	-	-	-
271	-	108	-	-	-
272	-	160	-	-	-
273	-	228	-	-	-

t	90%	95%	100%
180	0.7	0.4	-
190	1.0	0.8	-
200	1.6	1.2	-
210	1.9	1.4	0.6
220	2.9	1.8	0.9
230	4.7	2.1	1.3
235	5.9	-	-
240	9.1	-	-
245	23.8	2.5	1.8
248	70.8	-	-
250	162	3.9	2.4
251	461	-	-
260	-	6.1	3.0
270	-	9.5	3.9
280	-	15.8	5.1
285	-	28.2	-
290	-	62	7.0
292	-	95	-
294	-	157	-
295	-	205	-
300	-	-	10
305	-	-	11.9
310	-	-	14.5
315	-	-	17.7
320	-	-	23.1
325	-	-	31.0
330	-	-	43.8
331	-	-	51.2
332	-	-	60.0
333	-	-	76.5
334	-	-	179

Bogorodski, 1905

t	κ 100%	94.392%	73.224%	54.323%
348	6390	6620	7510	8300
351	6490	6710	7600	8430
354	6580	6800	7700	8540
357	6670	6890	7790	8650
360	6760	6990	7870	8750
363	6850	7090	7960	8850
t	**49.386%**	**14.493%**	**0%**	
348	8480	10450	11220	
351	8570	10480	11370	
354	8660	10550	11520	
357	8770	10670	11660	
360	8880	10880	11790	
363	9010	-	11920	

Goodwin and Mailey, 1908

t	κ(mhos)	t	κ(mhos)
0%		80 mol%	
318.7	1.022	322.4	0.6392
329.7	.073	332.7	.6849
349.8	.172	337.1	.6930
351.1	.184	352.4	.7399
363.3	.228	369.4	.8055
378.5	.299	388.7	.8684
379.7	.316	398.5	.9004
381.6	.318	449.0	1.055
397.8	.376		
404.3	.402		
411.4	.433		
417.9	.448		
420.1	.456	100 mol%	
428.9	.487	396.1	0.6563
446.3	.570	355.3	.7047
459.9	.595	364.5	.7179
477.6	.697	381.1	.7662
20 mol%		382.5	.7710
321.6	0.9119	387.5	.7887
399.1	1.224	403.4	.8291
454.1	.401	420.0	0.9000
50 mol%		429.0	.9179
		432.0	.9204
323.9	0.7625	434.2	.9250
364.2	.9262	474.4	1.039
367.4	.9313	491.0	.086
376.2	.9645	508.9	.131
387.6	.9981		
414.7	1.091		
433.5	.155		
442.0	.181		

Sandonnini, 1920

%	κ	%	κ
	550°		
100	6700	28.90	10300
92.62	7400	0	11700
54.84	8720		

Bergman and Chagin, 1940

%	κ 250°	275°	300°	325°
0	-	-	-	10640
10	-	-	892	9780
20	-	-	820	9040
30	-	6820	7640	8460
40	5420	6240	7060	7860
50	5020	5730	6520	7300
60	4720	5480	6250	7030
70	-	5260	6020	6790
80	-	-	5720	6460
90	-	-	-	6220
100	-	-	-	-

%	350°	375°	400°	425°
0	11600	12640	13640	14620
10	10680	11600	12500	13400
20	9900	10780	11600	12480
30	9370	10120	10960	11760
40	8660	9440	10240	11040
50	8080	8880	9660	10460
60	7300	8560	9310	10080
70	7520	8260	9070	9800
80	7200	7980	8740	9500
90	6960	7680	8420	9200
100	6660	7420	8180	8900

%	450°	475°	500°
10	15560	16600	17200
20	14280	15100	15960
30	13320	14160	14960
40	12580	13400	14210
50	11860	12650	13510
60	11240	12040	12840
70	10840	11600	12340
80	10240	11000	11780
90	9920	10690	11400
100	9640	10360	11070

Sodium nitrate ($NaNO_3$) + Rubidium nitrate ($RbNO_3$)

Pushin and Radoicic, 1937

mol%	f.t.	mol%	f.t.
0	306	55	178.5 E
10	291	60	185
20	203	70	200
30	240	80	220
40	214	90	260
50	190	100	314

Sodium nitrate ($NaNO_3$) + Thallium nitrate ($TlNO_3$)

van Eyk, 1900

mol%	f.t.	mol%	f.t.
100	206.1	58.9	204.1
98.1	203.5	46.5	228.2
96	198.4	37.4	242.5
84.7	176.4	25.1	263.7
76.5	163.2	15.5	277.4
71.7	174	0	308

Sodium nitrate ($NaNO_3$) + Silver nitrate ($AgNO_3$)

Hissink, 1899

mol%	f.t.	m.t.	mol%	tr.t.
100	208.6	208.6	100	159.8
92	211.4	210	98.93	154.2
84.94	215	212	97.55	148-149
80.54	217.2	214.8	95.85	139-140
78.1	222	215	95.50	138
70.3	228.4	216.5	95	138
63.8	234.8	217.5	91	138
52.7	259.4	237.6	80	138
41.4	272	257		
28	284	274		
0	308	308		

t	mol% mixed crystals I	mol% mixed crystals II
15	98.4	35.6
25	98.2	36.5
50	97.8	38.0

Retgers, 1889

%	d	%	d
6.05	2.346	28.99	2.595
15.28	2.425	37.16	2.738
22.08	2.518	45.95	2.872
26.63	2.567	52.32	2.982

Byrne, Fleming and Wetmore, 1952

density of melt. $g.cm^{-3} = a - b(t - 300)$

%	a	$b.10^5$	%	a	$b.10^5$
0	1.914	70	40	2.379	79
10	2.013	71	60	2.764	70
20	2.126	77	80	3.169	90
30	2.249	79	100	3.867	100

Bloom and Rhodes, 1956.

$n_D = a - bt$

mol%	t	a	$b.10_4$
20	310-380	1.504	1.0
40	290-360	1.535	0.7
60	270-350	1.576	0.6
80	235-330	1.632	0.8

Benrath, 1908

t	$\kappa.10^6$	t	$\kappa.10^6$	t	$\kappa.10^6$
		100 %			
180	3.2	200	14.8	206	43.1
185	5.0	202	19.8	207	60.3
190	7.3	204	27.5	208	99.2
195	9.6	205	33.7	209	258
		95 %			
180	4.2	200	18.4	206	45.1
185	5.9	202	22.8	207	59.2
190	8.7	204	30.4	208	82.9
195	12.8	205	35.8	209	125
		90 %			
180	5.1	200	23.3	206	63.9
185	7.5	202	30.4	207	93.7
190	9.0	204	41.4	208	164
195	13.9	205	49.2	209	317
		80 %			
180	6.0	202	18.9	212	63.9
190	9.4	209	27.9	214	80.8
195	10.9	210	34.6	215	187
200	13.3	211	43.1		
		75 %			
180	5.0	200	11.5	214	48.9
185	5.9	205	14.8	215	76.3
190	7.3	210	24.0	216	128
195	9.2	212	32.7		
		70 %			
180	4.0	200	8.6	214	28
185	4.6	205	11.0	215	30.9
190	5.3	210	23.8	216	110
195	7.0	212	25.2	217	290
		60 %			
180	3.3	202	12.2	216	48.6
185	3.9	204	23.7	217	66.5
190	4.8	212	29.5	218	98
195	6.3	214	34.9	219	140
200	8.7	215	38.2		
		50 %			
180	3.2	205	9.8	222	36.7
185	3.8	210	13.8	224	76.3
190	4.6	215	18.6	225	242
200	7.3	220	27.9		
		40 %			
180	2.9	225	15.1	239	109
190	4.3	230	20.2	240	144
200	5.5	235	42.8	241	180
210	8.1	237	62	242	182
220	10.7				
		30 %			
180	1.6	230	8.3	252	32.2
190	2.6	235	9.0	255	71.5
200	3.6	240	10.6	256	112
210	5.0	245	13.9	257	166
220	6.7	250	23.7		
		20 %			
180	0.8	230	4.3	265	23.5
190	1.2	239	5.2	270	43.2
200	1.6	245	7.2	272	61.0
210	2.3	252	8.8	274	86.2
220	3.0	260	12.5	275	212
		10 %			
180	0.6	239	4.9	280	43.2
190	0.8	245	6.3	282	49.9
200	1.2	260	8.4	284	59.0
210	1.8	265	11.1	286	73.0
220	2.8	270	17.0	287	94
230	3.8	275	25.5	288	154
		0 %			
210	0.4	260	3.6	300	24.8
220	0.7	270	5.9	305	60.2
230	1.1	280	7.8	306	76.5
239	1.5	290	14.0	307	142
245	2.4	295	18.8	308	200

Byrne, Fleming and Wetmore, 1952

wt%	mol%	κ 370°	κ 360°
0	0	12510	12090
10	5.27	12670	12240
20	11.12	12640	12230
30	17.66	12770	12350
40	25.01	12720	12300
60	42.88	12890	12470

mol %	κ 350°	340°	330°	320°
0	11650	11200	10760	10320
5.27	11800	11340	10880	10410
11.12	11820	11390	10940	10460
17.66	11940	11500	11040	10570
25.01	11890	11470	11030	10590
42.88	12040	11610	11180	10750
66.68	-	11800	11410	11000
100	-	-	11730	11730

mol %	310°	300°	290°
0	9870	-	-
5.27	9950	9480	9000
11.12	10010	9540	9070
17.66	10090	9600	9100
25.01	10130	9670	9200
42.88	10320	9870	9410
66.68	10590	10170	9730
100	10920	10510	10100

Aziz and Wetmore, 1952

ions transport 330°

mol%	f
5.069	0.806
11.552	.600
17.918	.315
24.353	.144

$f = (N_2 - N_2^{\circ})(n_1 + n_2)/Z$

where N_2 and N_2° are respective final and initial mole fractions of silver nitrate

n_1 and n_2 are respective equivalents of sodium and silver nitrate

Z is the number of faradays of charge

Sodium nitrate ($NaNO_3$) + Calcium nitrate (CaN_2O_6)

Menzies and Dutt, 1911

mol%	f.t.	mol%	f.t.
0	315.1	42.9	284.1
5.3	303.9	53.9	355.4
11.2	285.6	66.7	421.0
17.7	272.9	81.9	490.0
25.0	249.8	100	561.9
33.4	136.5		

Lehrman and Breslow, 1938

%	f.t.	E	%	f.t.	E
0	305.5	-	43.0	234	-
15.4	287	-	45.3	231	221.1
24.8	273	-	47.8	249	222.4
30.8	-	223.0	50.4	272	223.1
31.5	263	222.8	54.8	309	223.5
35	253	-	57.1	-	-
38.4	244	-	64.2	374	222.8
41.9	236	-	66.4	-	222.8

Protsenko and Bergman, 1950

%	f.t.	%	f.t.
100	561	29.8	232
66.7	454	29.1	236
60.0	424	27.4	240
53.9	394	25.0	245
48.2	365	21.2	253
46.0	347	17.7	263
42.9	330	14.3	271
40.9	312	11.2	278
38.0	298	8.2	287
36.1	278	5.3	294
33.4	257	2.6	303
31.6	240	0	308

Sodium nitrate ($NaNO_3$) + Strontium nitrate (SrN_2O_6)

Protsenko and Bergman, 1951

%	f.t.	%	f.t.
0	308	7.6	321
1.6	305	8.7	331
2.6	302	9.9	342
3.1	300	11.2	352
3.7	297	12.4	365
4.2	295	14.3	377
4.8	292	17.7	396
5.1	287 E	21.2	420
5.3	291	25.0	442
5.9	298	29.1	460
6.3	305	33.4	480
7.0	312		

Gromakov and Gromakova, 1953

mol%	f.t.	m.t.	
0	308	308	A
3	306	-	"
5	302	-	"
8	300	300	-
(9)	300	-	(1+1)
10	310	-	B
15	364	-	"
20	398	-	"
30	452	-	"
40	496	300	"
50	538	300	"
60	556	444	"
70	586	500	"
100	645	645	"

Sodium nitrate ($NaNO_3$) + Barium nitrate (BaN_2O_6)

Harkins and Clark, 1915

%	f.t.	%	f.t.
0	314.1	69.70	484.20
14.56	296.6	78.16	492.64
27.71	358.3	85.98	501.69
39.66	396.51	93.23	529.65
50.58	435.31	100	595.53
60.53	462.98		

Protsenko and Bergman, 1951

%-	f.t.	%	f.t.
0	308	14.3	364
1.3	305	16.0	384
2.6	302	17.7	392
3.9	300	21.2	406
5.3	297	25.0	421
6.7	294 E	29.1	440
7.1	297.5	33.4	456
8.2	312	38.0	474
9.6	330	42.9	486
11.2	345	48.2	502
12.7	353	53.9	517

Sodium nitrate ($NaNO_3$) + Lead nitrate (PbN_2O_6)

Guthrie, 1884

%	f.t.
0	305
42.84	268 E

Glass, Laybourn and Madgin, 1932

%	f.t.	%	f.t.
0.0	310.0	41.0	276.8
5.0	307.2	42.0	275.6
10.0	304.1	42.3	275.0 E
15.0	300.7	42.5	275.7
20.0	297.0	43.0	277.2
25.0	292.8	45.0	284.1
30.0	288.0	50.0	300.9
35.0	283.1	55.0	317.6
40.0	278.0	60.0	335.3

Laybourn and Madgin, 1932

%	transverse breaking strength (in Kg/cm_2)
60	0.53
55	.60
53	.54
50	.46
45	.12
42.3	.084
40	.105
35	.26
30	.27
25	.26
20	.22
15	.20
10	.18

Sodium nitrate ($NaNO_3$) + Cadmium nitrate (CdN_2O_6)

Protsenko, 1952 (fig.)

mol%	f.t.	mol%	f.t.
100	350	33.4	200
60	280	14.3	270
47.1	135 E	0	306

Popovskaya and Protsenko, 1954

	κ (mhos)			
mol%	180°	200°	220°	240°
60.0	-	-	-	-
53.9	-	0.032	0.056	0.088
51.0	0.021	.042	.073	.125
49.3	.025	.047	.083	.120
47.1	.028	.055	.093	.132
45.5	.031	.058	.097	.141
42.9	.043	.073	.115	.158
40.4	.045	.080	.125	.176
38.0	.045	.090	.136	.190
35.6	.061	.096	.148	.203
33.4	-	.111	.160	.219
31.8	-	-	0.180	0.243
29.1	-	-	-	.266
27.0	-	-	-	.284
25.0	-	-	-	.306

mol%	260°	280°	300°	320°
60.0	0.105	0.1495	0.195	0.245
53.9	.132	.1795	.230	.279
51.0	.154	.205	.257	.307
49.3	.167	.218	.269	.319
47.1	.180	.230	.284	.335
45.5	.187	.240	.292	.346
42.9	.208	.263	.315	.369
40.4	.227	.283	.336	.390
38.0	.243	.302	.360	.416
35.6	.259	.320	.382	.446
33.4	.278	.343	.406	.470
31.8	.307	.374	.442	.510
29.1	.324	.389	.460	-
27.0	.354	.424	.485	.564
25.0	.378	.449	.521	.592
21.2	.408	.488	.568	.646
17.7	.458	.537	.614	.692
11.2	-	.634	.735	.792
0	-	-	-	1.020

Potassium nitrate (KNO_3) + Rubidium nitrate ($RbNO_3$)

Pushin and Radoicic, 1937

mol%	f.t.	mol%	f.t.
0	337	60	291.5
10	322	70	290.5
20	313	80	293
30	307	90	301
40	299.5	100	311
50	294		

Protsenko and Popovskaya, 1953 and 1954

mol%	f.t.	mol%	f.t.
100	320	50	300
80	295	40	305
70	292	20	320
60	295	0	370

Stortenbeker, 1914

%	d	%	d
6.9	2.135	43.5	2.324
12.6	.150	46.5	.345
28.8	.226	47.9	.339
30.1	.250	48.3	.327
30.4	.222	48.3	.353
31.4	.242	50.6	.358
32.6	.252		

50.6 - 95.6 gap mixed crystals.

mol%	angle of optical axes 2E
0	10°51'
-	8°
4.8	6°3'
9	4°12'
22	2°32'
24.6	2°8'
35	1°45'
35.5	2°3'
36.3	2°33'
37	2°48'

Protsenko and Popovskaya, 1953 and 1954 (fig.)

			κ			
mol%	300°	320°	340°	360°	380°	400°
0	-	4220	4700	5160	5630	6080
90	-	4350	4850	5350	5810	6290
85	-	4380	4860	5340	5820	6300
80	3970	4460	4940	5440	5930	6420
75	4000	4490	5000	5500	6010	6500
70	4020	4540	5040	5550	6060	6560
67.5	4060	4590	5120	5600	6140	6660
65	4080	4600	5110	5640	6150	6740
62.5	4130	4640	5170	5700	6220	6790
60	4170	4700	5220	5760	6280	6860
55	4230	4760	5290	5820	6350	6990
50	-	4880	5410	5860	6470	7080
45	4380	4920	5480	6020	6560	7180
40	4440	4990	5560	6100	6640	7320
35	-	5090	5660	6210	6760	7430
30	-	5190	5750	6320	6870	7690
20	-	5420	5980	6560	7120	7870
15	-	-	6040	6640	7220	8210
10	-	-	6100	6700	7290	-
0	-	-	6350	6975	7598	-

van Eyk, 1899

Mixed crystals

mol%	tr.t.	mol%	tr.t.
100	79	100	72.8
92.6	82.5	95.6	75
80.6	86	85.4	77.5
76.5	86	75.3	78.4
69.3	86	64.5	78
59.4	86		

Protsenko and Shelomov, 1953 (fig.)

mol%	f.t.	mol%	f.t.
100	206	50	250
75	185	25	300
72	181 E	0	337

Potassium nitrate (KNO_3) + Thallium nitrate ($TlNO_3$)

van Eyk, 1905

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	206	205	50.8	242	219
88.6	197	190	36.5	270.5	-
77.6	188	182	25	292	270
67.6	191.8	182	12	317	307
60.6	217.5	182	0	339	-

mixed crystals

mol%	tr.t.$_1$	mol%	tr.t.$_2$
0	129.4	100	147
7	122	86	137.4
14	113	83.5	136.8
16.5	108.5	80	136.5
22.5	108.5	69.2	136
30.7	108.5	59.4	136-105
31.8	108.5-133	49.5	136-105
40.5	108.5-133	39.9	136-105
45.2	108.5-133	30.6	136-105
49.7	108.5-133	20.6	105
60.5	108.5-133	0	125.6
68.7	133		
71.4	133		
77.6	135.5		
84.1	136.5		
85.3	139.5		
94.2	143.3		
100			

$_1$= heating curve

$_2$= cooling curve

Potassium nitrate (KNO_3) + Silver nitrate ($AgNO_3$)

Ussov, 1904

mol%	f.t.	mol%	f.t.
100	209	57	122.8
90	192	55.5	128
85	183	55	131
82.5	178.4	51.5	151
79.6	171.8	40	200.5
76.1	164.7	30	241
74	160.2	22	275.5
72.1	156.6	10	311.2
69.6	150	5.5	325.4
64.5	136.3	0	339
63.1	132.9		
62.3	131	60.2	131.8
60.2	125.1	58.3	133 (1+1)
58	120	55.6	133.9

mol%	E	mol%	tr.t.
90	119	85	131
80	120	80	131
70	"	76.1	130.5
67	"	74	131
62.5	"	72.5	"
60	120.2	64.5	"
58.1	120	62.3	"
58	"	57	130.9
57	"	51.5	134.3
55	"	48	"
51.5	"	40	"
30	119	30	"
20	119.5		

Doucet and LeDuc, 1953 (fig.)

mol%	f.t.	mol%	f.t.
0	334.17	55.6	125 E
20	235	80	170
40	165	100	208.9

Protsenko, 1953 (fig.)

mol%	f.t.
0	180
25	155
38	130
41	135
48	134

Lifshits, 1956

%	f.t.	%	f.t.
100	208	54.5	134 (1+1)
62	131 E	0	337

Bloom and Rhodes, 1956.

$n_D = a - bt$

mol%	t	a	b.10_4
30	250-350	1.538	2.0
55	150-300	1.588	1.9
58	190-300	1.593	1.9
70	170-300	1.614	1.4

Cowen and Axon, 1956

mol %	κ 200°	250°	300°
100	6000	8400	10500
85	5200	7600	9600
75	4800	7000	9000
65	4300	6400	8500
50	3700	5700	7400
35	3000	4900	6700
22	2500	4300	6000
0	2000	3500	5000

Zakharyevskii and Permyakova, 1956.

mol%	e	mol%	e
	320°		
14.3	+93.0	40.0	+20.4
25.0	+57.3	40.0	+20.1
28.6	+48.0	41.7	+16.9
33.3	+36.5	44.26	+11.7
38.19	+24.6	58.16	+17.4
39.66	+18.6	64.95	-31.7
39.66	+16.0	73.56	-52.6
		84.78	-89.3

mol %	e	mol%	e
	250°		
30.8	+37.2	54.19	-6.5
40.0	+18.1	54.19	-8.1
44.26	+10.7	64.95	-28.8
46.10	+ 6.8		

mol%	e	mol%	e
180°		150°	
46.10	+5.5	46.10	+3.2
54.19	-5.7	51.2	-0.2
58.16	-8.9	52.65	-1.3
		54.19	-3.3
140°		135°	
45.0	+0.8	51.77	+1.0
52.65	-0.8	51.20	+0.8
52.16	-0.3	49.00	-1.6
51.20	0.0	48.10	-1.3
47.68	0.0		

Potassium nitrate (KNO_3) + Calcium nitrate (CaN_2O_6)

Menzies and Dutt, 1911

%	f.t.	%	f.t.
0	346.3	53.9	345.3
5.3	318.5	66.7	450.0
11.2	295.2	81.9	508.4
17.7	262.2	100	221.9
42.9	263.3		

Protsenko and Bergman, 1950

%	f.t.	%	f.t.
100	561	48	165
75	403	45	174
70	357	42.5	192
65	321	40	204
62.5	283	35	230
60	260	30	257
57.5	236	25	277
55	208	20	297
52.5	171	15	309
51	146	10	320
50	153	5	329
49	157	0	337

Rostkovski, 1930

mol%	f.t.	(4+1)	E	tr.t.
0	337	-	-	127
2.6	328	-	-	-
5.3	321	174	-	127
8.2	308	-	-	-
11.2	292	-	-	-
14.3	272	174	-	-
17.6	260	"	-	-
19.5	233	"	-	-
21.2	224	"	145	-
23.1	211	-	-	-
24.3	200	-	-	-
26.6	186	174	145	-
28.3	174	-	-	-
29.8	162	-	-	-
31.6	154	-	-	-
32.5	145	-	-	-
35.6	174	-	-	-
38	183	-	-	-
39,9	210	-	-	-
40.9	230	-	-	-
41.9	250	-	-	-
43.9	270	-	-	-
46	290	-	-	-
48.2	310	-	-	-
50.4	329	-	145	-
53.9	368	-	-	-
60.0	402	-	145	-
100	-	-	-	-

E : 34.2 mol% 145° (4+1) : 28.7 mol% 174°
tr.t. : 28.7 mol% (incongr.)

Natsvlishvili and Bergman, 1943

mol%	κ (mhos)					
	425°	400°	375°	350°	325°	300°
0	8930	8180	7350	6660	-	-
5.3	8300	7550	6820	6100	5380	-
8.2	8000	7280	6620	5880	5220	-
11.2	7800	7050	6400	5700	5100	-
14.3	7480	6750	6030	5350	4830	3900
16.3	7300	6600	5900	5250	4600	3750
20.5	7250	6600	5820	5120	4200	3550
21.2	7250	6580	5750	5100	4400	3750
22.0	6930	6170	5400	4680	4250	3500
22.7	6900	6200	5430	4750	4170	3400
24.3	6700	6050	5320	4600	3950	3300
25.9	6500	5900	5200	4520	3800	3100
27.4	6350	5700	5050	4400	3680	3000
29.1	6250	5600	4900	4270	3600	2900
31.6	6000	5380	4750	4100	3400	2800
33.4	5870	5220	4600	3900	3300	2700
35.2	5500	4950	4300	3650	3100	2550
42.9	4900	4300	3650	3150	2550	1950
48.2	4550	3950	3300	2850	2100	1500
53.8	4100	3550	3000	2550	-	-
60.0	3900	2800	-	-	-	-

mol%	κ (mhos)						
	275°	250°	225°	200°	175°	150°	145°
14.3	3200	2680	-	-	-	-	-
16.3	3100	2400	-	-	-	-	-
20.5	2800	2100	1500	-	-	-	-
21.2	3100	2400	1900	-	-	-	-
22.0	2750	2100	1500	-	-	-	-
22.7	2700	1950	1400	-	-	-	-
24.3	2600	1850	1300	-	-	-	-
25.9	2450	1800	1250	-	-	-	-
27.4	2400	1750	1200	850	-	-	-
29.1	2300	1700	1100	850	400	150	-
31.6	2170	1500	1020	750	350	150	-
33.4	2050	1350	970	650	300	120	70
35.2	1920	1200	850	500	270	100	-
42.9	1300	600	-	-	-	-	-

Potassium nitrate (KNO_3) + Strontium nitrate (SrN_2O_6)

Protsenko and Bergman, 1951

mol%	f.t.	mol%	f.t.
0	337	17.0	304
2.6	326	17.7	308
3.9	320	19.5	324
5.3	316	21.2	338
6.7	311	23.1	354
8.2	306	25.0	369
9.6	301	27.0	384
11.2	295	29.1	400
12.7	291	33.4	424
14.3	278	38.0	417
15.0	284	42.9	472
15.7	290	48.2	494
16.3	294	53.9	513

E : 25% 277.5° Higher than 70% = decomposition.

Gromakov and Gromakova, 1953

mol%	f.t.	m.t.	
0	336	-	
5	324	-	
10	296	247	
15	279	266	
(18)	266	-	E
20	297	266	
25	356	-	
35	418	-	
45	480	314	
55	513	399	
70	562	-	
100	645	645	

Potassium nitrate (KNO_3) + Barium nitrate (BaN_2O_6)

Harkins and Clark, 1915

%	f.t.	%	f.t.
0	346.3	65.96	474.31
12.56	333.8	75.05	493.87
24.41	311.9	83.75	495.32
35.65	350.0	92.01	514.42
46.25	380.03	100	595.53
56.35	433.9		

Protsenko and Bergman, 1951

mol%	f.t.	mol%	f.t.
0	337	19.5	338
1.3	331.5	21.2	350
2.6	328	23.1	364
3.9	322	25.0	380
5.9	314	27.0	390
6.7	308	29.1	400.5
8.2	305	33.3	423
9.6	302	38.0	443
11.2	292	42.9	463
12.7	287 (E)	48.2	477
14.3	301.5	53.9	494
16.0	312	60.0	511
17.7	326		

Potassium nitrate (KNO_3) + Lead nitrate (PbN_2O_6)

Guthrie, 1884

%	f.t.	%	f.t.
0	320	46.82	207
10	300	50	210
20	285	60	238
30	268	70	335
40	246		

Glass, Laybourn and Madgin, 1932

%	f.t.	%	f.t.
0.0	340.0	48.0	224.1
5.0	335.0	49.0	219.7
10.0	328.2	49.5	217.8 E
15.0	320.8	50.0	219.2
20.0	311.9	51.0	222.4
25.0	301.7	52.0	226.4
30.0	290.4	53.0	230.2
35.0	275.2	55.0	240.7
40.0	257.6	60.0	271.0
45.0	238.0	65.0	303.4
47.0	228.0		

Laybourn and Madgin, 1932

%	*	%	*
60	1.32	30	1.55
55	1.06	25	1.09
50	0.70	20	0.72
45	1.06	15	0.55
40	1.27	10	0.42
35	1.37		

* : transverse breaking strength in Kg/cm_2

Potassium nitrate (KNO_3) + Cadmium nitrate (CdN_2O_6)

Protsenko, 1952 (fig.)

mol%	f.t.
100-67	decomposition
66.7	250
45.5	168 E
33.4	199.5 (2+1)
25.9	175 E
14.3	270
0	337

Popovskaya and Protsenko, 1955

mol%	d 180°	200°	220°	240°
65.0	-	2.630	2.612	2.593
62.0	2.510	2.590	2.571	2.552
60.0	2.585	2.565	2.546	2.526
57.5	-	2.539	2.519	2.500
55.0	-	2.504	2.485	2.467
52.5	-	2.479	2.460	2.440
50.0	-	2.448	2.429	2.410
47.5	-	2.422	2.403	2.384
45.0	-	2.398	2.378	2.359
42.0	2.384	2.368	2.350	2.331
40.0	-	2.346	2.328	2.309
37.5	-	2.319	2.301	2.283
35.0	-	-	2.279	2.252

mol%	d 260°	280°	300°
70.0	2.646	2.627	2.608
65.0	2.573	2.554	2.534
62.0	2.533	2.514	2.495
60.0	2.507	2.488	2.468
57.5	2.481	2.462	2.443
55.0	2.448	2.429	2.410
52.5	2.421	2.402	2.383
50.0	2.392	2.373	2.355
47.5	2.365	2.346	2.327
45.0	2.340	2.322	2.303
42.0	2.312	2.294	2.276
40.0	2.291	2.273	2.254
37.5	2.266	2.247	2.229
35.0	2.243	2.224	2.207
30.0	2.196	2.180	2.160
25.0	-	2.140	2.115

Protsenko and Popovskaya, 1954

mol%	κ 180°	200°	220°	240°
60	100	230	500	715
54	220	440	740	1025
51	300	530	830	1160
48	340	620	980	1355
45	400	710	1090	1485
43	470	790	1170	1586
40.5	-	875	1300	1680
38	700	1015	1420	1825
35	-	1095	1505	1960
33.5	-	1160	1610	2060
31	-	1800	1650	2120
29	890	1280	1770	2240
27	-	1330	1830	2300
25	990	1400	1880	2350
21	-	-	2060	2530
	260°	280°	300°	320°
60	1040	1420	1835	2300
54	1390	1815	2290	2800
51	1545	1985	2450	3020
48	1770	2220	2710	3300
45	1910	2380	2915	3400
43	2032	2530	3060	3530
40.5	2200	2650	3120	3700
38	2365	2740	3260	2840
35	2400	2890	3410	3910
33.5	2510	2970	3465	4060
31	2650	3110	3600	4200
29	2710	3180	3700	4280
27	2850	3300	3900	4430
25	2900	3400	3930	4520
21	3030	3500	4100	4800
17.5	3310	3800	4380	4900
14.5	-	3900	4500	5090
11	-	-	4600	5300
5.5	-	-	-	5600

Rubidium nitrate ($RbNO_3$) + Cesium nitrate ($CsNO_3$)

Protsenko and Belova, 1955

Solid solutions

Rubidium nitrate ($RbNO_3$) + Silver nitrate ($AgNO_3$)

Protsenko and Kiparenko, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	310	56 E	122
20	210	80	155
40 E	128	100	209
50			

Palkin, 1926

mol%	f.t.	E	tr.t.
100	208	-	160
95	195	128	159
90	182	"	160.5
85	169	"	-
75	146	"	-
72.5	140	"	-
70	134	"	-
67.5	128	"	-
65	131.5	"	-
60	135.5	"	-
55	138	-	-
50	139.5	-	-
45	138	136	-
42.5	137	"	-
40	136	"	-
37.5	141	"	-
35	149	"	-
30	167	"	-
25	190	141	162
20	212	"	"
15	235	"	"
10	257	"	"
5	-	"	"
0	306	-	-

Rubidium nitrate ($RbNO_3$) + Strontium nitrate (SrN_2O_6)

Plyushchev, Markina and Shklover, 1956.

mol%	f.t.	tr.t. 1	tr.t. 2	E	tr.t. 3
0.0	314	291	219	-	164
1.0	301	291	220	199	166
2.0	288	-	217	198	168
3.0	275	-	216	196	166
5.0	266	-	219	200	162
7.5	240	-	215	202	161
10.0	229	-	220	205	164
12.5	220	-	220	196	160
15.0	210	-	-	200	164
17.5	205	-	-	199	164
20.0	200	-	-	200	162
22.5	218	-	-	200	162
25.0	250	-	-	201	160
27.5	275	-	-	205	161
30.0	292	-	-	203	165
32.5	305	-	-	201	163
35.0	327	-	-	199	165
37.5	357	-	-	204	163
40.0	373	-	-	200	163
45.0	420	-	-	202	161
50.0	448	-	-	202	165
55.0	483	-	-	202	165
60.0	576	-	-	198	162
65.0	540	-	-	201	-
70.0	565	-	-	197	-
75.0	580	-	-	200	-
80.0	590	-	-	201	-
90.0	615	-	-	203	-
100.0	645	-	-	-	-

Rubidium nitrate ($RbNO_3$) + Calcium nitrate (CaN_2O_6)

Protsenko and Belova, 1955

(1 + 1)

Rubidium nitrate ($RbNO_3$) + Barium nitrate (BaN_2O_6)

Plyushchev, Markina and Shklover, 1956

mol%	f.t.	tr.t. 1	E	tr.t. 2	tr.t. 3
0.0	314	291	-	219	164
1.0	299	291	235	217	162
2.0	295	290	240	217	167
5.0	275	-	240	219	162
10.0	253	-	230	212	162
15.0	250	-	231	217	160
20.0	244	-	237	214	164
25.0	240	-	234	221	170
27.0	238	-	235	214	163
29.0	235	-	235	219	164
30.0	240	-	235	216	164
35.0	355	-	239	213	158
40.0	402	-	235	219	159
45.0	452	-	239	222	166
50.0	473	-	240	220	164
55.0	494	-	234	221	167
60.0	502	-	235	222	-
65.0	512	-	230	217	-
70.0	520	-	237	218	-
75.0	530	-	234	214	-
80.0	543	-	233	-	-
85.0	554	-	232	-	-
90.0	573	-	237	-	-
95.0	581	-	234	-	-
100.0	592	-	-	-	-

Protsenko and Popovskaya, 1953 (fig.)

mol %	f.t.
100	350
66.7	310
42.9	190
37.5	167 E_1
33.3	199.5
22.7	152 E_2
11.2	240
0	310

(2+1)

Protsenko and Popovskaya, 1954

mol%	$\varkappa.10^6$ 180°	200°	220°	240°
55	-	-	-	1260
62.5	-	-	1020	1360
60	480	750	1060	1420
57.5	540	790	1120	1490
55	600	890	1210	1600
52.5	590	880	1240	1620
50	610	990	1250	1670
47.5	620	990	1290	1700
45	650	1000	1290	1780
42.5	720	1000	1390	1750
40	760	1020	1390	1810
37.5	790	1060	1420	1840
35	800	1100	1450	1860
30	850	1150	1540	1960
20	-	-	-	2040
	260°	280°	300°	320°
80	-	-	-	1990
75	-	-	1840	2260
70	-	1770	2000	2600
67.5	1420	1820	2240	2650
65	1620	2020	2400	2840
62.5	1650	2160	2500	2900
60	1800	2200	2570	3000
57.5	1860	2280	2680	3120
55.0	2000	2410	2820	3200
52.5	2020	2450	2850	3250
50	2060	2420	2870	3380
47.5	2100	2550	2950	3430
45	2170	2620	3070	3500
42.5	2220	2640	3080	3520
40	2210	2650	3100	3620
37.5	2220	2680	3100	3680
35	2260	2750	3200	3850
30	2350	2850	3250	4000
20	2460	2970	3400	5200
10	-	-	-	4380

Cesium nitrate ($CsNO_3$) + Thallium nitrate ($TlNO_3$)

Protsenko and Rubleva, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	413	50	285
25	450	75	230
40	315	100	206

Cesium nitrate ($CsNO_3$) + Silver nitrate ($AgNO_3$)

Palkin, 1928

mol%	f.t.	E	tr.t.
100	208	-	159
97.5	203.5	168	"
95	198.5	"	"
92.5	192.5	168.5	"
90	187	168.5	"
85	175	"	"
82.5	168.5	"	-
80	170	"	-
77.5	171	"	-
75	171.5	-	-
72.5	170	-	-
70	168.5	163	-
67.5	163	"	-
65	166	"	-
62.5	169	"	-
60	171	-	-
55	172	-	-
52.5	173	173	-
50	197	172.5	151
45	238	172	151
40	264	172	151
35	288	"	"
25	331	"	"
20	349	171.5	"
10	380	171	"
0	407	-	-
	(1+3)	(1+1)	

Cesium nitrate ($CsNO_3$) + Calcium nitrate (CaN_2O_6)

Protsenko and Belova, 1955

(1 + 1)

Cesium nitrate ($CsNO_3$) + Strontium nitrate (SrN_2O_6)

Plynshchev, Markina and Shklover, 1956

mol%	f.t.	E	tr.t.
0.0	414	-	154
2.5	401	270	155
5.0	392	272	155
7.5	386	275	157
10.0	371	273	152
12.5	355	276	153
15.0	347	276	153
17.5	333	277	155
20.0	309	272	153
22.5	287	278	153
25.0	275	275	155
27.5	295	275	157
30.0	321	275	157
32.5	368	275	155
35.0	396	277	154
37.5	403	273	155
40.0	424	276	154
45.0	458	278	153
50.0	495	280	152
55.0	520	270	155
60.0	544	280	155
65.0	565	275	156
70.0	585	275	-
80.0	608	274	-
85.0	625	274	-
90.0	630	275	-
100.0	645	-	-

Cesium nitrate ($CsNO_3$) + Barium nitrate (BaN_2O_6)

Plynshchev, Markina and Shklover, 1956.

mol%	f.t.	E	tr.t.
0.0	414	-	154
5.0	388	310	158
10.0	345	311	154
13.0	310	310	150
15.0	326	310	154
17.0	331	310	155
20.0	348	310	157
25.0	370	311	157
30.0	385	310	154
35.0	412	310	152
40.0	435	312	152
45.0	463	317	160
50.0	480	312	150
55.0	488	305	151
60.0	507	307	157
65.0	530	313	149
70.0	535	305	-
75.0	550	308	-
80.0	563	305	-
85.0	575	306	-
90.0	580	310	-
95.0	587	305	-
100.0	592	-	-

Cesium nitrate ($CsNO_3$) + Lead nitrate (PbN_2O_6)

Pushin and Radoicic, 1937

mol%	m.t.	mol%	m.t.
0	418	32	176
10	362	40	260
20	292	50	317
30	200		

Cesium nitrate ($CsNO_3$) + Cadmium nitrate (CdN_2O_6)

Protsenko and Popovskaya, 1954

mol%	f.t.	mol%	f.t.
100	350	28.3	166 E
66.7	305	25.0	220
42.9	200	11.2	350
40	162.5 E	0	410
33.3	178 (2+1)		

Protsenko and Rubleva, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	413	33	178
5	375	39.4 E	162.5
14	345	43	190
25	235	54	255
29.4 E	166	67	305

(2+1)

Protsenko and Popovskaya, 1954

mol%	κ(mhos) 160°	180°	200°	220°
48.2	-	-	0.0640	0.092
42.9	0.032	0.052	.0745	.104
40.4	.0355	.0565	.0795	.108
38.0	-	.060	.0830	.110
35.6	-	.063	.0860	.115
33.3	-	.064	.0890	.117
31.2	-	.065	.0900	.118
29.1	-	.066	.0895	.119
27.0	-	-	.0910	.120
25.0	-	-	-	.123
23.1	-	-	-	-

mol%	κ(mhos) 240°	260°	280°	300°	320°
66.7	-	-	-	0.153	0.190
60	-	-	0.148	.1835	.221
53.9	0.103	0.136	.1695	.2045	.240
48.2	.121	.153	.187	.2235	.258
42.9	.1335	.165	.1975	.2300	.2625
40.4	.136	.1685	.201	.234	.268
38.0	.141	.1725	.206	.239	.271
35.6	.145	.1765	.209	.242	.275
33.3	.146	.178	.210	.245	.279
31.2	.148	.180	.212	.247	.281
29.1	.1485	.180	.213	.249	-
27.0	.150	.182	.215	.250	.283
25.0	.153	.186	.220	.255	.288
23.1	.1555	.189	.222	.257	.2905
21.2	-	-	.226	.262	.300

Thallium nitrate ($TlNO_3$) + Silver nitrate ($AgNO_3$)

Sandonnini, 1920

%	κ	%	κ
	250°		
0	4360	50.0	5800
10.0	4650	75.0	6950
25.0	5120	100.0	8120

van Eyk, 1899

mol%	f.t.	mol%	f.t.
100	208.5	52	81.6
94	196	50	82.8
90	188.5	48	81.2
87	183	47	85
86	181	41.5	100.6
78.5	161.5	40.5	105
70.5	140	37.5	112
68	133	30.7	129
62	116	22.5	149
54	91	14	173.5
53	85	0	200

Thallium nitrate ($TlNO_3$) + Lead nitrate (PbN_2O_6)

Glass, Laybourn and Madgin, 1932

%	f.t.	%	f.t.
0.0	206.2	16.0	185.5
2.5	203.1	17.0	192.6
5.0	197.8	20.0	214.2
7.5	192.0	25.0	247.5
10.0	186.0	30.0	279.1
12.0	181.3	35.0	306.9
13.0	179.1	40.0	335.6
14.0	170.9	45.0	357.1
14.7	175.5 E	50.0	378.0
15.0	177.0		

Thallium nitrate ($TlNO_3$) + Cadmium nitrate (CdN_2O_6)

Protsenko and Rubleva, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	206	33	105
5	190	43	205
14	152	54	270
27.4 E	95	67	305

(2+1)

Protsenko and Popovskaya, 1954

%	100°	120°	140°	160°	180°
			κ(mhos)		
47.5	-	-	-	0.0825	0.1165
50	-	0.0345	0.060	.092	.130
52.5	-	.042	.0715	.104	.142
55	0.026	.049	.080	.1185	.156
57.5	.0295	.054	.086	.124	.165
60	.033	.059	.093	.1325	.174
62.5	-	-	.100	.136	.177
65	-	-	.106	.144	.188
70	-	-	.122	.161	.206
75	-	-	-	.172	.220
80	-	-	-	-	.239
85	-	-	-	-	.250

%	200°	220°	240°	260°	280°	300°
30		-	-	-	0.220	0.261
35	-	-	0.171	0.215	.2595	.304
40	-	0.164	.205	.249	.295	.340
42.5	0.136	.177	.220	.264	.309	.354
45	.1485	.190	.234	.278	.323	.369
47.5	.156	.199	.244	.288	.334	.380
50	.170	.213	.257	.303	.349	.396
52.5	.184	.228	.275	.322	.369	.415
55	.200	.247	.294	.341	.387	.434
57.5	.210	.258	.306	.353	.401	.448
60	.219	.266	.315	.369	.410	.457
62.5	.229	-	.320	.375	-	-
65	.235	.283	.330	.380	.427	.474
70	.254	.303	.351	.400	.448	.495
75	.268	.318	.368	.417	.467	.517
80	.289	.339	.389	.4385	.488	.537
85	.302	.352	.402	.451	.501	.551
90	.317	.367	.4165	.467	.516	.566
100	-	.396	.446	.498	.548	.599

mol%	f.t.	mol%	f.t.
100	350	43	95
80	300	40	100
60	220	20	170
55	101	0	225
50	101		

(2+1)

Silver nitrate ($AgNO_3$) + Lead nitrate (PbN_2O_6)

Glass, Laybourn and Madgin, 1932

%	f.t.	%	f.t.
0.0	209.5	12.0	230.9
2.5	208.2	15.0	249.8
5.0	205.4	20.0	278.7
7.5	205.2	25.0	302.2
8.5	200.2 E	30.0	323.6
9.0	206.2	35.0	342.5
10.0	215.7	40.0	359.3
11.0	223.3		

Silver nitrate ($AgNO_3$) + Cadmium nitrate (CdN_2O_6)

Protsenko, 1953 (fig.)

%	f.t.	%	f.t.
75	233	44	151
55	148	0	185
50	155		

Popovskaya and Protsenko, 1955

mol%	160°	170°	190°	210°
		d		
60.0	-	-	-	3.509
57.5	-	3.566	3.545	.523
55.0	3.596	.585	.564	.541
52.5	.616	.605	.583	.561
50.0	.633	.622	.600	.578
47.5	.651	.640	.618	.596
45.0	.669	.658	.635	.613
42.5	.686	.674	.651	.629
40.0	-	.693	.668	.647
37.5	-	-	.694	.672
35.0	-	-	.712	.688
25.0	-	-	.785	.762
20.0	-	-	-	.794

mol%	230°	250°	270°	290°
65.0	-	3.430	3.410	-
60.0	3.487	.466	.444	3.427
57.5	.502	.480	.459	.437
55.0	.520	.498	.477	.456
52.0	.538	.516	.495	.473
50.0	.550	.533	.510	.488
47.5	.574	.552	.530	.508
45.0	.590	.568	.546	.524
42.5	.607	.584	.562	.540
40.0	.624	.601	.579	.556
37.5	.649	.625	.602	.580
35.0	.666	.643	.620	.597
25.0	.739	.716	.693	.669
20.0	.771	.748	.724	.702
0	.928	.905	.882	.858

Barium nitrate (BaN_2O_6) + Lead nitrate (PbN_2O_6)

Tammann and Krings, 1923

%	Q mix.	
5-6°		
100.0	25.6	by dissolving
67.0	30.3	" "
35.1	33.5	" "
18.5	36.8	" "
0.0	39.8	" "
67.0	30.1	heated at 340° during 24h.
35.1	33.3	" " " "
18.5	36.8	" " " "
54.0	31.5	mecanic mixture

Vegard and Dale, 1928

mol%	α (in A°)
100	7.836
90.3	.867
85.74	.875
74.99	.904
62.97	.940
46.82	.986
39.23	8.010
38.12	.010
28.5	.039
17.5	.072
0	.110

Bismuth nitrate (BiN_3O_9) + Didymium nitrate (DiN_3O_9)

Bodman, 1901

mixed crystals + (6 aq.)		mixed crystals + (5 aq.)	
%	d	%	d
48.6	2.762	88.51	2.830
51.8	.486	83.60	.754
66.5	.477	81.57	.723
71.1-	.398	100.00	.700
70.1	.377	0.00	.284
71.3	.381		
77.4	.375		
100.0	.347		
0.0	.249		

Lithium metaphosphate ($LiPO_3$) + Potassium metaphosphate (KPO_3)

Bergman and Sholokhovich, 1953

mol%	f.t.	mol%	f.t.	
100	798	64	536	
96	774	60	544	
92	744	57.5	551	(1+1)
88	720	50	562	
84	694	45	550	
80	666	40	532	
76	641	35	528	
72	609	30	543	
68	582	25	566	

E_1 : 63% 528° E_2 : 36% 518°

Sodium metaphosphate ($NaPO_3$) + Potassium metaphosphate (KPO_3)

Morey, 1954

wt%	m.t.	
0	627.6	
10	606	
20	582	
27.85[a]	563	
30	555	
31	552	
34	551	
36.66	550	
40	548.4	
45	547.9	
50	547.1	
50.5	547	
55	583.1	(3+1)
70	676	
80	721.5	
100	813	

Tammann and Ruppelt, 1931

mol%	f.t.	E	mol%	f.t.	E
0	650	-	60	665	550
10	640	630	70	725	545
20	605	550	80	755	545
30	620	"	90	775	715
40	675	"	100	810	-
50	650	"			

Sodium metaphosphate ($NaPO_3$)
+ Calcium metaphosphate (CaP_2O_6)

Morey, 1952

%	f.t.	%	f.t.
0	627.6	40.4	734
-	625 E	44.8	728
4.6	661	49.2	739
9.4	696	70.0	834
25.8	734	89.8	927
32.8 (2+1)	738	100.0	977
E_1 : 1 %		E_2 : 46 %	725°

Sodium metaarsenate ($NaAsO_3$)
+ Potassium metaarsenate ($KAsO_3$)

Amadori, 1911

%	f.t.	E	min.	tr.t.
0	615	-	-	-
10	580	-	-	-
15	550	-	-	-
20	526	487	40	-
25	515	492	80	-
30	508	495	140	-
33.3	502	"	190	-
40	-	496	222	-
50	498	-	-	-
60	504	-	-	-
65	512	-	-	-
70	526	-	-	-
80	550	-	-	-
85	576	-	-	-
90	602	-	-	368
95	624	-	-	412
97.5	638	-	-	428
100	650	-	-	450

Sodium orthoarsenate (Na_3AsO_4)
+ Potassium orthoarsenate (K_3AsO_4)

Amadori, 1911

%	f.t.	sat.t. (mixed crystals)	tr.t.
0	1260	-	410
5	1238	-	388
10	1220	-	370
15	1205	-	370
20	1190	442	-
30	1164	454	-
40	1156	464	-
50	1175	464	-
60	1192	462	-
70	1225	455	-
80	1255	450	-
85	1265	438	-
90	1280	-	-
95	1290	-	-
100	1310	-	-

Lithium vanadate (Li_3VO_4) + Lead vanadate (PbV_2O_6)

Bukhalov and Aleshkina, 1953 (fig.)

mol%	m.t.	mol%	m.t.
0	1230	71.5 E	736
-	-	80	810
30	1070	100	952

Sodium pyrophosphate ($Na_4P_2O_7$)
+ Potassium pyrophosphate ($K_4P_2O_7$)

Amadori, 1913

mol%	f.t.	tr.t. I	tr.t. II
0	994	395	-
2.43	981	248	-
4.07	973	180	-
5.71	966	-	-
8.20	953	-	-
12.43	943	-	-
16.76	922	-	-
21.15	897	-	-
25.66	887	-	-
34.94	875	-	-
44.60	882	-	505
50.00	892	-	505
54.70	908	-	460
65.26	943	-	370
76.32	988	-	-
82.02	1016	-	-
87.80	1036	-	-
92.67	1055	-	-
96.30	1075	-	-
100	1090	-	-

Morey, Boyd Jr., and al., 1955

%	f.t.	%	f.t.
0	989	50.40	894
17.98	930	55.39	900
26.50	906	63.65	920
34.74	898	68.60	937
37.95	891	70.68	951
38.32	890	71.00	956
41.13	890	80.17	989
42.70	894	100	1104
44.26	889		

Belyaev and Sholokhovich , 1953 (fig.)

mol%	f.t.	mol%	f.t.
100	1095	42	882 min.
75	975	25	920
50	895	0	992

Sodium pyroarsenate ($Na_4As_2O_7$) + Potassium pyroarsenate ($K_4As_2O_7$)

Amadori, 1924

%	f.t.	E	min.	tr.t.I	tr.t.II
0	850	-	-	688	-
5	825	-	-	652	-
10	813	-	-	612	-
15	792	-	-	576	-
20	788	-	-	540	-
25	777	528	110	528	-
30	768	532	100	-	-
35	766	545	70	-	425
40	768	550	60	600	428
45	774	550	50	612	438
50	776	542	30	620	450
54.17	790	-	-	620	450
60	809	-	-	612	450
65	822	-	-	592	446
70	842	-	-	580	440
75	854	-	-	558	-
80	890	380	30	540	-
85	902	384	40	520	-
90	932	390	70	-	-
95	963	390	100	-	-
100	996	-	-	400	-

Sodium tetraborate ($Na_2B_4O_7$) + Calcium tetraborate (CaB_4O_7)

Jilenko and Sverchkov, 1939

mol%	f.t.	mol%	f.t.
100	960	40	870
95	976	37	852
90	985	35	834
85	988	30	790
80	990	28	770
75	986	25	738
70	980	22	706
65	968	21	glass
60	958	20	"
55	940	10	"
50	924	0	"
45	904		

Lithium carbonate (Li_2CO_3) + Sodium carbonate (Na_2CO_3)

Skaliks, 1929

%	f.t.	E	min
0	735	-	-
10	718	490	30
20	680	494	102
30	640	505	166
40	613	510	238
50	549	510	280
58.9	-	514	- (1+1)
60	-	510	-
65.9	570	"	198
80	708	502	101
100	860	-	-

Lithium carbonate (Li_2CO_3) + Potassium carbonate (K_2CO_3)

Le Chatelier, 1894 and 1897

%	f.t.	%	f.t.
100	860	65.16	515
94.98	777	60.96	505
90.45	720	53.41	492
88.21	682	53.41	492
80.64	590	48.41	525
74.14	515	35.85	600
79.16	492	26.99	638
74.14	500	15.61	673
70.25	505	0	710

Lithium carbonate (Li_2CO_3) + Calcium carbonate ($CaCO_3$)

Skaliks, 1929

%	f.t.	E	min
0	735	-	-
5	733	656	-
10	732	"	32
16.7	715	650	69
30	703	662	-
40	670	663	290
50	720	663	254
55	765	660	210
65	-	662	167
70	-	657	236
74.7	990	654	-

Sodium carbonate (Na_2CO_3) + Potassium carbonate (K_2CO_3)

Le Chatelier, 1894

mol%	f.t.	%	mol%	f.t.	%
100	860	100	35	700	41.25
80	770	83.91	21	740	25.74
62	715	68.02	0	820	0
45	690	51.61			

Amadori, 1913

mol%	f.t.	tr.t.	mol%	f.t.	tr.t.
0	854	(430)	50.00	709	490
7.85	813	-	60.54	740	475
16.94	770	-	75.42	793	-
24.74	726	-	87.35	844	-
36.66	706	480	93.58	868	-
43.40	704	490	100	896	405

Niggli, 1913, 1916 and 1919

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	896	-	46	712	-
79	819	-	44	712	-
72	791	757	33	721	-
63	760	737	27	739	712
53	725	-	15	774	753
47	712	-	0	860	-

Belyaev and Sholokhovich, 1952

mol %	f.t.	mol %	f.t.	mol %	f.t.
100	896	65	757	30	721
95	888	60	739	25	738
90	865	55	730	20	757
85	841	50	720	15	785
80	821	45	711	10	812
75	798	40	704	5	843
70	775	35	707	0	857

E = 60 % 704°

Makarov and Shulgina, 1940

mol%	n		mol%	n	
	650°	550°		650°	550°
0.0	1.547	1.547	50.0	1.536	1.540
5.0	.536	.540	65.0	.536	.540
10.0	.536	.540	97.0	.536	.540
33.0	.536	.540	100.0	.539	.539
	400°	300°		400°	300°
0.0	1.547	1.547	50.0	1.550	1.548
5.0	.533	.5288	65.0	.533	.528
20.0	.535	-	97.0	.533	.528
33.0	.539	.535	100.0	.539	.539

Makarov and Shulgina, 1940

mol%	f.t.	m.t.	tr.t.	
			I	II
0	-	-	356	-
"	-	-	486	-
"	-	-	620	-
2.5	-	-	320	348
"	-	-	450	478
5.0	-	-	294	340
"	-	-	466	466
"	-	-	620	-
10.0	792	788	252	310
"	-	-	352	442
15.0	-	-	436	-
20.0	754	744	420	-
25.0	-	-	412	-
"	-	-	490	-
30.0	728	720	490	-
33.0	-	-	490	-
"	-	-	620	-
35.0	-	-	490	-
40.0	712	706	490	-
40.0	-	-	618	-
45.0	736	718	500	-
"	-	-	620	-
50.0	-	-	490	-
"	-	-	620	-
60.0	760	744	490	-
"	-	-	620	-
65.0	782	756	490	-
"	-	-	620	-
70.0	-	-	200	-
"	-	-	486	-
"	-	-	620	-
75.0	-	-	264	-
"	-	-	440	-
"	-	-	476	-
"	-	-	630	-
77.0	823	800	200	-
"	-	-	470	-
80.0	-	-	220	-
"	-	-	264	-
"	-	-	465	-
"	-	-	618	-
83.0	848	830	258	-
"	-	-	436	-
85.0	-	-	254	-
"	-	-	420	-
"	-	-	620	-
90.0	-	-	210	-
"	-	-	334	-
"	-	-	460	-
"	-	-	618	-
95.0	-	-	226	278
"	-	-	380	-
"	-	-	518 sic	-
97.0	-	-	210	240
"	-	-	380	408
100.0	-	-	250	-
"	-	-	428	-
"	-	-	622	-

Sodium carbonate (Na_2CO_3) + Calcium carbonate ($CaCO_3$)

Niggli, 1916

mol%	f.t.	m.t.	tr.t.
0	860	-	-
7	874	868	-
11	873	868	-
20	863	854	-
28	840	772	-
30	831	786	-
36	805	786	-
40	801	787	-
47	810	784	-
50	813	-	813
52	-	-	815
54	-	-	818
64	-	-	812
66	-	-	814
67	-	-	817
74	-	-	814

Eitel, 1925 (fig.)

%	f.t.	m.t.	tr.t.
0	864	-	-
10	880	872	-
18	889	889	-
30	840	775	-
41	775	775	-
50	815	775	815
60	885	-	815
70	965	-	815
80	1060	-	815
90	1210	-	815
100	1340	-	815

Sodium carbonate (Na_2CO_3) + Barium carbonate ($BaCO_3$)

Sackur, 1912

m	D. f.t.
0.59	-17
0.75	23
1.08	32
1.31	37

Potassium carbonate (K_2CO_3) + Calcium carbonate ($CaCO_3$)

Niggli, 1916

mol%	f.t.	tr.t.	E
0	895	-	895
4	-	890-885	-
10	-	881-872	-
14	-	867-853	740
20	-	842	751
25	-	821	751
38	-	769	753
42	-	787	752
44	-	801	752
50	811	-	811
50	812	-	812
54	-	-	800
58	-	821	800

Calcium carbonate ($CaCO_3$) + Strontium carbonate ($SrCO_3$)

Faivre and Chaudron 1948

Stability of mixed crystals.

Calcium carbonate ($CaCO_3$) + Barium carbonate ($BaCO_3$)

Faivre and Chaudron, 1948

Stability of mixed crystals .

Kordes, 1926

mol %	f.t.
0	1259
48	1139
100	1740

Calcium carbonate ($CaCO_3$) + Ferrous carbonate ($FeCO_3$)

Diesel, 1911

%	d	%	d
	21°		
0	2.57	44.59	2.694
12.56	.607	46.70	.70
19.08	.609	65.51	.733
38.57	.685		

Strontium carbonate ($SrCO_3$) + Barium carbonate ($BaCO_3$)

Faivre and Chaudron, 1948

Stability of mixed crystals .

Sodium calcium carbonate ($Na_2CaC_2O_6$) + Potassium calcium carbonate ($K_2CaC_2O_6$)

Niggli, 1919

mol%	f.t.	mol%	f.t.
100	814	50	810
76	805	30	813
70	806	22	815
60	808	0	817
54	810		

Mixed crystals.

Lithium metasilicate (Li_2SiO_3) + Sodium metasilicate (Na_2SiO_3)

Wallace, 1909

%	f.t.	m.t.	%	f.t.	m.t.
100	1018	1018	35	984	984
90	975	951	30	1022	1022
80	881	869	25	1064	1064
70	831	817	20	1090	1090
60	786	786	10	1156	1156
50	813	800	0	1168	1168
40	841	841			

Kracek, 1939

%	f.t.	m.t.	n_D (quenched glass)
100	1089	1089	1.517
90	1022	975	.527
85	988	919	.528
80	955	882	.529
76.5	930	863	.529
73.5	905	851	.530
70	885	848	.532
67.5	869	848	.534
66	861	847	.534
64	851	845	.536
64	851	845	.537
62	846	845	.536
61.2	846	845	.538
60.7	-	845	.539
60.7	848	845	.539
60	852	845	.538
59	860	846	-
57.6	872	847	.539
53.6	872	846	.539
57.6	872	847	.540
57	877	847	.540
50	931	846	.542
40	998	846	.545
30	1072	-	-
20	1125	-	-
15	1147	860	-
10	1169	970	-
5	1185	1077	-
0	1201	1201	1.557

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	1200	54 E	824
24	1072	62	851
33	986	75	940
48	884	100	1088

(1+2) (2+1)

Lithium metasilicate (Li_2SiO_3)
+ Potassium metasilicate (K_2SiO_3)

Wallace, 1909

%	f.t.	m.t.	%	f.t.	m.t.
100	vitreous		40	947	933
90	"		30	999	972
80	"		20	1112	1078
70	"		10	1152	1128
60	"		0	1168	1168
50	"				

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	1200	53 E	774
20	1065	75	850
27	986	86 E	814
40	890	100	960
48	795		

(1+3) (3+2) - incongruent

Lithium metasilicate (Li_2SiO_3)
+ Magnesium metasilicate ($MgSiO_3$)

Wallace, 1909

%	f.t.	m.t.	E	min.
0	1168	1168	-	-
10	1122	1066	-	-
20	1111	1025	-	-
30	1091	1002	-	-
40	946	-	-	-
50	-	-	873	100
55	986	-	895	85
60	1088	-	886	70
65	1158	-	-	-
70	1328	-	-	-
80	1430	-	-	-
90	1430	-	-	-
100	1449	1449	-	-

Lithium metasilicate (Li_2SiO_3)
+ Calcium metasilicate ($CaSiO_3$)

Wallace, 1909

%	f.t.	m.t.	E	min.	d
0	1168	1168	-	-	2.61
10	1147	1096	-	-	.64
20	1121	1075	-	-	.69
30	1083	-	-	-	.72
40	1043	-	969	90	.76
50	-	-	987	140	.80
60	1056	-	979	80	.80
70	1114	-	980	40	.84
80	1252	-	973	20	.85
90	1315	-	-	-	.88
100	1502	1502	-	-	.91

Bergman and Bichkova, 1955 and Nesterova, 1955

mol%	f.t.	mol%	f.t.
0	1201	43	1034
25	1106	46	1027
28	1094	49	1050
31	1082	52	1072
34	1072	55	1100
37	1060	57.5	1126
40	1046	100	1540

E : 1024°

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	1200	45 E	1024
25	1150	57	1160
40	1085		

Lithium metasilicate (Li_2SiO_3)
+ Barium metasilicate ($BaSiO_3$)

Wallace, 1909

%	f.t.	m.t.	E	min.
0	1168	1168	-	-
10	1141	1094	-	-
20	1120	1062	-	-
30	1088	-	-	-
40	1063	-	857	10
50	1011	-	881	15
60	955	-	835	20
70	912	-	860	25
75	900	-	868	30
80	985	-	877	70
85	1143	-	894	40
90	1238	-	883	25
95	1359	-	-	-
100	1490	1490	-	-

Next Page

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	m.t.	mol%	m.t.
65	1085	30	1090
55 E	960	0	1200
48	986		

Bychkova and Bergman, 1956.

mol%	f.t.	mol%	f.t.
30	1092	50	980
32.5	1080	52.5	972
35	1070	55	960 E
37.5	1058	57.5	985 tr.t.
40	1045	60	1015
42.5	1033	62.5	1045
45	1017	65	1082

Wallace, 1909

%	d	%	d
0	2.61	60	3.50
10	.70	70	.77
20	.80	80	4.02
30	.91	90	.19
40	3.17	100	.44
50	.38		

Lithium metasilicate ($LiSiO_3$) + Zinc metasilicate ($ZnSiO_3$)

van Klooster, 1910

%	f.t.	%	f.t.
0	1188	60	-
10	1184	70	-
20	1169	80	-
30	1130	90	1353
40	-	100	1419
50	-		

Lithium metasilicate (Li_2SiO_3) + Aluminum metasilicate ($Al_2Si_3O_9$)

Ballo and Dittler, 1912

mol%	f.t.	m.t.	E		d C	d E
0	1180	1180	-		2.4893	-
10	1105	1035	-		.4622	-
20	1025	970	-		.4280	-
30	960	940	940		.3268	-
31.82	960	938	938	E	.3882	-
33.33	965	965	-	(2+1)	.3880	-
35.00	930	915	915	E	-	2.3822
37.50	1070	920	920	-	2.3531	.3800
40.00	1250	-	-	-	.3650	.3785
50.00	1275	1275	1275	(1+1)	.3127	.3111
58	1255	1200	1200	-	.3222	.4286
65	1240	1210	1210	-	-	-
70	-	-	-	-	-	2.4823

Sodium metasilicate (Na_2SiO_3) + Potassium metasilicate (K_2SiO_3)

Sholokhovich and Barkova, 1955

mol %	f.t.	mol %	f.t.
0	1082	63.6	768
20	954	64	772
30	874	66	778
35	860	67.5	780
36	844	69	785
40	828	70	788
42	814	73	795
45	800	75	795
48	792	79	781
51 E_1	784	82	753
54	788	85	772
57.5	793	88	800
60	793	91	834
61	788	94	863
63	777	100	970

E_2 = 768° (1+3) (2+3)

Previous Page

Sodium metasilicate (Na_2SiO_3) + Magnesium metasilicate ($MgSiO_3$)

Wallace, 1909

%	f.t.	%	f.t.
0	1018	60	vitreous
10	921	70	"
20	vitreous	80	"
30	"	90	1435
40	"	100	1549
50	"		

Sodium metasilicate (Na_2SiO_3) + Calcium metasilicate ($CaSiO_3$)

Kultascheff, 1903

%	f.t.	%	f.t.
0	1007	50	1118
10	970	55	1138
20	938	60	1160
25	1022	65	1146
30	1060	70	1142
35	1090	80	1128
40	1090	85	1150
42.5	1098	100	more than 1400

(2+3)

Wallace, 1909

%	f.t.	%	f.t.	E
0	1018	65	1165	-
10	986	70	1132	-
20	932	75	-	1153
30	1024	77.5	1258	1110
38.8	1086	80	1394	-
45	1091	85	1449	-
50	1143	90	1450	-
58.8	1175	100	1502	-

(2+3)

Ginsberg and Nicogossian, 1924

%	f.t.	%	f.t.
0	1093	60	1255
10	1078	70	1230
20	1042	75	1258
30	1080	80	1310
40	1118	90	1390
50	1206	100	1512

(2+3) Mixed crystals.

Morey and Bowen, 1925

Two additive compounds

Sodium metasilicate (Na_2SiO_3) + Strontium metasilicate ($SrSiO_3$)

Wallace, 1909

%	f.t.	%	f.t.
0	1018	50	960
10	889	60	986
20	875	70	1180
30	914	80	1312
35	907	90	1454
40	930	100	1529

Sodium metasilicate (Na_2SiO_3) + Barium metasilicate ($BaSiO_3$)

Wallace, 1909

%	f.t.	m.t.
0	1018	1018
10	1015	1015
20	995	995
30	947	918
40	908	908
50	929	914
60	961	936
70	1010	967
75	1101	1079
80	1187	1147
90	1378	1378
100	1490	1490

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	1088	33 E	929
15	1015	36	975
25	945		

Magnesium metasilicate ($MgSiO_3$)
+ Calcium metasilicate ($CaSiO_3$)

Allen and White, 1909

%	f.t.	E	d
100	1510	-	2.912
92	-	-	.965
95	1484	1343	-
90	1467	1344	2.947
80	1404	1345	3.046
76	-	1350	-
74	-	1348	-
72	-	1348	-
70	-	1349	3.111
69	-	1350	-
68	-	1349	-
64	-	1348	-
60	1374	1347	3.201
55	1380	1343	.229
53.7	1380	-	.236
52.2	1381	-	.246
50	1381	-	.245
45	1378	-	.241
40	-	-	.229
35	1378	-	.221
30	-	1376	.205
28	1409	1375	.196
27	1410	1373	-
25	1416	1372	3.194
20	1452	1374	.198
10	1488	1375	.188
5	1512	1387	.181
0	1524	-	.183

(1+1)

Tillotson, 1918

vol%	n_D	d	vol%	n_D	d
0	1.5801	2.758	58.75	1.6105	2.858
4.76	.5823	.777	62.80	.6122	.872
9.54	.5853	.781	72.99	.6174	.881
28.92	.5960	.823	84.60	.6224	.891
38.76	.6008	.835	94.75	.6261	.899
52.35	.6073	.854	100	.6280	.904

Magnesium metasilicate ($MgSiO_3$)
+ Ferrous metasilicate ($FeSiO_3$)

Bowen and Shairer, 1935 and
Shairer and Bowen, 1942 (fig.)

%	f.t.	m.t.	tr.t. I	tr.t. II
0	1560	1560	-	1145
20	1530	1530	-	1130
30	1510	1420	-	1120
40	1475	1400	1370	1100
61	1400E	1400	1300	1040
70	1460	-	1310	1020
88	1510	-	-	960
100	1580	-	-	-

Magnesium metasilicate ($MgSiO_3$)
+ Manganese metasilicate ($MnSiO_3$)

Lebedev, 1911

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	1210	1210	50	1328.6	1304
91.4	1223	1216	41.55	1368	-
87.5	1237.5	-	25	1437	1395
82.8	1256	1222	16.7	1450	-
75	1270	1229	0	1535	1535
58.4	1313	1280			

Calcium metasilicate ($CaSiO_3$)
+ Barium metasilicate ($BaSiO_3$)

Lebedev, 1911

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	1438	1438	34.5	1040	-
93.75	1337	-	25	1279	1220
81.25	1228	-	18.7	1393	-
75	1172	1130	12.5	1455	1383
61.6	1080	1024	0	1512	1512

Calcium metasilicate ($CaSiO_3$)
+ Ferrous metasilicate ($FeSiO_3$)

Konstantinov and Selivanov, 1912

%	f.t.	m.t.	%	f.t.	m.t.
100	1150,1160	-	0	1510	-
88.8	1096	1025	95	1125	-
77.3	-	1030	80	1054	1025
69.1	1100	"	60	1130	-
67.3	1112	1020	44.2	1185	-
59.5	1150	-	35	"	-
58.8	1156	1040	30	1200	-
51.9	1175	-	20	1350	1200
27.5	1357	1220	5	1475	-
27.4	1335	1144			
17.9	1424	1223			
1.6	1505	1140			

E : 75% 1025°

Bowen, Shairer and Posnjak, 1933; Shairer and Bowen, 1942 (fig.)

%	f.t. I	f.t. II	tr.t. 1	tr.t. 2	m.t. II	tr.t. 3
0	1545	-	-	-	-	-
5	1530	-	-	1230	-	-
10	1510	-	-	1280	-	-
20	1470	-	-	1225	-	-
30	1390	-	1280	1175	-	-
40	1320	-	1280	1160	1160	880
53	1210	-	-	-	1120	960
63	1160E	1160	-	-	1110	950
70	1260	1140	-	-	1105	940
75	1330	1105	-	-	1105	980
80	1360	1105E	-	-	-	-
90	1470	1170	-	-	-	-
100	1470	1190	-	-	-	-

Calcium metasilicate ($CaSiO_3$) + Manganese metasilicate ($MnSiO_3$)

Ginsberg, 1908

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	1512.0	-	37.5	1259	-
88.8	1467.5	-	27.5	1232	-
78	1410.0	-	22.9	1208	-
67.4	1373.0	1303	12.8	1184	1184
57	1343.0	1272	0.2	1198	-
51.9	1319	-	0	1218	-
63.9	1285	-			

Kallenberg, 1914

mol%	f.t.	n	Birefringence
100	1380	1.636	positive
90	1360	.639	"
80	1340	.648	negative
70	1320	.658	"
60	1290	.666	"
50	1270	.678	"
40	1250	.687	"
30	1220	.698	"
20	1180	-	"
10	1150	-	"
0	1180	1.714	positive

Ginsberg, 1908

%	d	%	d
0	2.919	70	3.219
15	.992	85	.302
40	3.080	90	.313
60	.180	100	.350

Kurnakov and Jemchuznii, 1908

mol%	hardness	mol%	hardness
	room temperature		
0	0.041	70	0.058
15	.045	85	.061
40	.055	100	.048

Barium metasilicate ($BaSiO_3$) + Manganese metasilicate ($MnSiO_3$)

Lebedev, 1911

mol%	f.t.
97	1188
94	1173
87.5	1137
10	1363
0	1438

Zinc metasilicate ($ZnSiO_3$) + Cadmium metasilicate ($CdSiO_3$)

van Klooster, 1910

%	f.t.	m.t.	%	f.t.	m.t.
0	1419	-	60	1147	-
10	1381	-	70	1052	-
20	1356	-	80	1082	-
30	1276	-	90	1144	1136
40	1220	-	100	1155	-
50	1174	1164			

Manganese metasilicate ($MgSiO_3$) + Ferrous metasilicate ($FeSiO_3$)

Ussov, 1904

%	f.t.	%	f.t.
0	1134	58	1164
14	1138	77	1174
33	1141	95	1166
40	1149	100	1215

Lithium titanate (Li_2TiO_3) + Potassium titanate (K_2TiO_3)

Belyaev and Sigida, 1956.

mol%	f.t.	mol%	f.t.
100	826	88	759
98	798	86	805
97	784 tr.t.	84	836
96	784	82	866
94	783	80	907
92	774	78	938
90	764	76	950
88.5	750 E	74	989

Lithium titanate (Li_2TiO_3) + Sodium titanate (Na_2TiO_3)

Belyaev and Sigida, 1956.

mol%	f.t.	mol%	f.t.
100	1030	81.5	992 E
95	1026	80	1015
90	1020	75	1053
75	1005	70	1097

Sodium titanate (Na_2TiO_3) + Potassium titanate (K_2TiO_3)

Nisioka, 1934

mol%	f.t.	tr.t.	E	min.
0	973	-	-	-
5	959	-	-	-
10	922	865	762	10
20	$L_1 + L_2$	897	766	15
30	-	892	768	30
40	-	894	780	"
50	-	894	780	45
60	863	-	778	65
70	828	-	765	90
80	778	-	"	110
87	789	-	777	50
90	796	-	776	"
95	808	-	778	20
100	810	-	-	-

Sholokhovich, 1955

mol%	f.t.	mol%	f.t.
100	826	45	906
95	812	40	906
90	798	35	904
85	798	30	910
80	830	25	934
75	852	20	958
70	873	15	966 (1+1)
65	886	10	996
60	890	5	1012
55	900	0	1025
50	906		

E_1 : 31.5% 902° E_2 : 86.5% 784°

Potassium titanate (K_2TiO_3) + Lead titanate ($PbTiO_3$)

Belyaev, Sholokhovich and Barkova, 1956.

mol%	f.t.	mol%	f.t.
0	826	25	817
10	810	35	830 (2+1)
19	786 E	39	815 E

Strontium titanate ($SrTiO_3$) + Barium titanate ($BaTiO_3$)

Smolenskii and Isupov, 1954 and 1956 (fig.)

mol%	tr.t. I	tr.t. II	tr.t. III
100	380	285	210
90	360	270	200
80	335	250	190
70	310	235	180
60	280	210	170
50	240	190	-
40	210	170	-
30	170	135	-
20	125	-	-
10	80	-	-

Smolenskii, 1956, fig.

mol%	tr.t. I	tr.t. II	tr.t. III
100	390	280	215
90	370	270	200
80	340	250	190
70	300	230	180
60	280	215	175
50	240	200	-
40	220	170	-
30	170	135	-
20	120	-	-
10	80	-	-

Khodakov, Sholokhovich and al., 1956.

%	t	d	η	ε
		23°		
100	120	6.01	2.40	1550
95	94	5.98	2.38	1200
90	88	5.94	2.33	1500
85	58	5.92	2.38	2700
50	-48	5.54	2.37	1550
0			2.35	

t = Curie point

Barium titanate ($BaTiO_3$) + Lead titanate ($PbTiO_3$)

Bergstein, 1955

Seignett electric properties of solid solutions.

Barium stannate ($BaSnO_3$) + Lead stannate ($PbSnO_3$)

Smolenskii, 1956 (fig.)

mol%	tr.t.	mol%	tr.t.
78	0	50	- 95
70	- 55	40	-130
60	- 65	30	-155

Lithium sulfate(Li_2SO_4) + Sodium sulfate(Na_2SO_4)

Le Chatelier, 1897

mol %	f.t.	mol %	f.t.
0	830	30	620
10	750	100	860
20	680		

Nacken, 1918

%	f.t.	tr.t.	E	min.
0	843	573	-	-
12.54	772	531	468	150
24.39	689	471	"	315
35.62	634	486	"	225
46.25	596	505	"	75
51.36	601	514	"	45
56.35	612	520	-	-
61.20	624	518	507	-
65.94	629	503	-	-
66.88	637	496	-	-
69.21	651	491	-	-
72.65	673	489	210	66
75.07	689	486	209	120
83.77	752	466	212	225
91.42	812	427	"	315
96.60	846	220	"	"
100	883	234	-	-

Akopov and Bergman, 1954

mol %	f.t.	mol %	f.t.	mol %	f.t.
100	884	58	636	36	608
90	828	55	628	33	628
85	796	50	626	30	644
80	768	48	620	25	676
75	700	45	616	20	712
70	696	42	611	15	744
65	660	40	608	10	782
62	644	38	604	0	856

(1+1) (1+2)

E_1 : 62.5 mol% 637° E_2 : 48 mol% 620°

E_3 : 36.5 mol% 596°

Bakumskaya and Bergman, 1956

mol %	f.t.	mol %	f.t.	mol %	f.t.
0	856	42.5	596	60	630
10	782	45	602	62	636 tr.t.
15	742	47.5	608	65	662
25	664	48	608 tr.t.	70	696
32.5	608	50	614	75	728
35	592	52.5	620	79	758 tr.t.
36.5	584 E	55	626	80	766
37.5	584	57.5	628	85	799
40	590				

Lithium sulfate (Li_2SO_4) + Potassium sulfate (K_2SO_4)

Nacken, 1918

%	f.t.	E	min.	$tr.t_1$	min.	$tr.t_2$	min.
0	843	-	-	573	-	-	-
7.69	758	-	-	564	315	-	-
14.96	651	535	90	565	105	420	45
21.84	556	"	210	-	-	"	75
28.35	535	"	270	-	-	421	"
34.54	613	541	210	-	-	422	"
40.42	648	538	150	-	-	"	90
46.02	677	540	135	-	-	423	135
51.35	699	536	90	-	-	429	165
56.43	714	-	-	430	-	420	195
61.29	716	-	-	-	-	435	-
65.93	710	696	80	-	-	"	180
70.37	698	698	195	579	-	"	135
78.69	799	"	150	594	30	434	105
86.36	893	"	60	595	60	"	90
93.44	992	"	45	"	90	432	45
100.00	1076	-	-	"	-	-	-

Dombrovskaya, 1933 (fig.)

mol %	f.t.	mol %	f.t.
100	1076	40	690
80	920	20	534 E
60	698 E	0	845
50	716 (1+1)		

Bergman, Kislova and Korobka, 1954

mol %	f.t.	mol %	f.t.	mol %	f.t.
0	858	20.5	532 E	60	713
5	760	21	536	60.5	712
10	645	25	588	61	720
12	592	35	694	70	810
12.5	574	45	732	75	864
13	572	50	736	80	908
15	561	55	728	100	1069
19	540				

Akopov and Bergman, 1954

mol %	f.t.	mol %	f.t.	mol %	f.t.
0	856	26	600	46	730
5	750	28	625	50	732
10	650	30	648	55	726
15	556	32	666	60	716
18	542	35	688	65	770
20	536	40	702	70	825
22	542	42	722	100	1069
24	574				

E_1 : 19 mol % 532° E_2 : 60.5 mol % 712°

(1+1) : 732° (congruent)

(2+1) : 23.5 mol % 550° (incongruent)

I - II Li_2SO_4 : 17 mol % 572°

Bergman and Kislova, 1955 and
Bergman, Kislova and Posipaiko, 1954 and 1955

mol %	f.t.	mol %	f.t.	mol %	f.t.
0	856	26	600	46	730
5	752	28	625	50	732
10	650	30	648	55	726
15	556	32	666	60	716
18	542	35	688	65	770
20	536	38	702	70	825
22	542	42	722	100	1069
24	574				

E_1 : 19 mol % 532° E_2 : 60.5 mol % 712°

(1 + 1) 732° (2 + 1) incongruent

Bergman and Bakumskaya, 1956

mol %	f.t.	mol %	f.t.
0	856	25	600
5	762	30	656
10	656	35	692
12.5	598	40	712
15	558	45	722
17.5	544	50	724 (1+1)
18.5	536 E	55	718
19	538	60	700 E
20	540	60	702
21	544	67.5	788
22	548 tr.t.(2+1)	75	860
22.5	554	80	908

Benrath and Drekopf, 1921

t	$\varkappa \cdot 10^6$	t	$\varkappa \cdot 10^6$
	0 mol %		
347	0	630	58880
374	0	655	64570
400	0	679	67610
426	0	704	74130
452	11.2	728	75860
478	33.1	752	83180
504	120	776	85110
530	380	799	89130
555	1122	823	93330
580	53700	846	257000
605	57540		
	5 mol %		
347	0.933	452	66.1
374	4.17	478	166
400	11.5	504	501
426	26.9	530	1349
	10 mol %		
347	1.27	452	55.0
374	3.89	478	141
400	8.51	504	407
426	20.0	530	2138
	20 mol %		
400	13.2	478	110
426	23.4	504	251
452	42.7	530	776
	30 mol %		
347	1.32	452	41.7
374	3.89	478	93.3
400	8.51	504	257
426	17.8	530	891
	40 mol %		
374	1.15	478	72.4
400	4.27	504	159
426	10.5	530	501
452	27.5		
	45 mol %		
426	10.0	504	148
452	25.1	530	380
478	61.7		
	50 mol %		
504	4.47	630	204
530	12.6	655	490
555	25.1	679	977
580	44.7	704	2290
605	89.1		
	60 mol %		
452	1.51	605	112
478	3.80	630	282
504	6.17	655	617
530	11.0	679	1510
555	19.1	704	3980
	70 mol %		
478	3.39	605	145
504	6.31	630	398
530	12.0	655	891
555	24.0	679	2190
580	43.7	704	5750
	80 mol %		
504	20.9	630	832
530	28.8	655	1660
555	52.5	679	3800
580	110	704	9120
605	398		
	90 mol %		
478	2.82	605	288
504	8.13	630	501
530	14.5	655	1200
555	28.2	679	2510
580	50.1	704	5750
	100 mol %		
555	3.16	728	33.9
580	4.47	752	47.9
605	6.31	776	69.2
630	8.91	799	100
655	12.6	823	126
679	17.4	846	174
704	24.0		

Lithium sulfate (Li_2SO_4) + Cesium sulfate (Cs_2SO_4)

Dergunov, 1951

mol%	f.t.	mol%	f.t.
100	1020	50	738 (1+1)
95	987	45	734
90	943	40	724
85	892	35	710
80	840	30	686
75	792	25	644
72.5	765	24.5	630
70	738	22.5	633
69	680 E	20	637
67.5	714	15	632
65	690	10.5	620 E
62	696	10	637
60	712	5	782
58	720	0	856
55	730		

Lithium sulfate (Li_2SO_4) + Silver sulfate (Ag_2SO_4)

Nacken, 1907

%	f.t.	tr.t.	(2+3) decomposes	E	min.
0.00	843	573	-	-	-
12.78	806	552	-	-	-
23.93	770	538	418	372	45
33.33	728	524	419	375	60
41.45	688	514	420	370	75
43.18	639	512	"	377	90
54.83	598	509	"	"	105
60.40	580	505	418	380	120
65.37	567	494	422	381	150
73.90	568	471	420	381	160
80.94	572	-	"	"	180
84.02	574	-	406	380	185
86.85	"	-	404	378	195
89.47	"	-	393	380	210
91.89	577	-	390	379	240
96.22	620	-	382	382	270
98.17	638	-	391	"	165
100.00	651	-	412	-	-

Lithium sulfate (Li_2SO_4) + Calcium sulfate ($CaSO_4$)

Le Chatelier, 1894

mol %	f.t.
0	830
10	750
20	675

Müller, 1910

%	f.t.	E	min.	tr.t.	min.
0	843	-	-	572	337
3.68	817	-	-	571	322
7.39	790	677	30	570	307
10.82	763	689	52	"	300
14	736	694	90	"	285
17.91	707	699	135	571	277
21.35	700	700	180	572	270
23.61	705	698	165	571	262
29.19	763	700	143	572	240
34.64	825	700	120	"	229
45.19	913	699	82	573	195
55.29	989	695	52	"	165
64.97	1054	676	37	572	135
74.26	1125	670	22	572	105

Plyushchev and Komisarova, 1952

mol%	f.t.	E	tr.t.
0.0	852	-	575
1.0	840	693	"
2.5	830	696	572
5.0	806	692	"
7.5	780	695	576
10.0	758	696	574
12.5	730	694	576
15.0	705	696	575
16.5	695	695	573
18.0	700	"	574
20.0	704	"	576
25.0	746	694	572
30.0	780	"	576
35.0	818	696	"
40.0	866	"	"
45.0	905	692	573
50.0	936	695	576
55.0	986	698	575
60.0	1023	693	574
65.0	1069	695	575
70.0	1120	"	-
75.0	1170	693	-
78.0	1205	696	1205
80.0	1233	695	1203
82.0	1251	"	1205
100.0	(1450)	-	"

Golubeva and Bergman, 1954

mol%	f.t.
100	1450
15	706 E
0	852

Golubeva and Bergman, 1955

E : 22 % 722°

Lithium sulfate (Li_2SO_4) + Strontium sulfate ($SrSO_4$)

Calcagni and Marotta, 1912

%	f.t.	E	min.	tr.t.	min.
0	856	-	-	585	40
3	841	-	-	"	-
5	831	731	5	"	37
7	816	736	7	"	-
10	811	746	15	"	36
15	791	751	20	582	35
20	766	746	22	587	34
22	753	741	24	585	32
25	-	746	-	587	30
30	816	"	22	585	28
35	881	747	20	"	-
40	925	751	18	"	25
45	961	746	17	"	-
50	994	"	-	587	-
55	1036	"	12	585	-
60	1081	751	12	582	16
65	1116	736	-	"	-
70	1131	731	-	"	-
75	1136	736	-	577	-
100	1225	-	-	-	-

Lithium sulfate (Li_2SO_4) + Barium sulfate ($BaSO_4$)

Calcagni and Marotta, 1912

%	f.t.	E	min.	tr.t.	min.
0	856	-	-	585	40
2	846	760	-	-	-
5	826	"	6	585	38
10	801	765	12	"	-
15	781	760	16	"	-
18	-	"	13	582	-
20	771	755	-	"	33
22	791	"	22	"	-
25	821	770	-	"	32
30	861	765	19	585	28
35	906	760	-	584	27
40	938	"	16	582	25
45	987	"	16	"	-
50	1033	765	-	"	21
55	1071	760	13	577	18
60	1111	750	12	582	-
65	1141	755	11	578	-
70	1136	750	-	579	-
75	1141	755	-	577	-
100	1350	-	-	-	-

Lithium sulfate (Li_2SO_4) + Lead sulfate ($PbSO_4$)

Calcagni and Marotta, 1912

%	f.t.	E	min.	tr.t. I	min.	tr.t. II
0	856	-	-	585	18	-
5	841	638	2	"	17	-
10	816	"	3	"	15	-
15	786	"	4	"	-	841
20	759	"	5	"	14	-
25	729	"	7	"	12	861
30	706	"	-	"	-	836
35	676	"	-	"	-	856
40	656	"	11	"	9	861
42	649	"	12	"	9	856
45	646	"	13	"	-	-
47	-	"	14	575	-	856
50	659	"	-	585	-	841
55	696	643	8	580	-	-
60	726	638	7	585	-	856
65	766	648	6	"	-	836
70	799	638	-	"	-	836
75	826	"	5	"	5	-
80	841	"	-	"	4	-
100	1005	-	-	-	-	856

Lithium sulfate (Li_2SO_4) + Zinc sulfate ($ZnSO_4$)

Evseyev and Bergman, 1951

mol%	f.t.	
0	850	I
-	654	"
35.5	618	"
36.7	610	"
37.5	600	"
39.0	581	I + II tr.t.
39.5	572	"
42.0	566	II
43.5	562	"
45.5	548	"
49.5	520	"
51.5	510	II + B E
52.0	520	B
53.0	530	"
54.0	542	"
55.0	556	"
57.3	576	"
60.5	612	"
63.3	626	"
67.5	647	"
70.0	657	"

Lithium sulfate (Li_2SO_4) + Cadmium sulfate ($CdSO_4$)

Calcagni and Marotta, 1913

%	f.t.	E	tr.t.
0	856	-	585
5	832	-	575
10	811	551	570
15	786	"	575
20	756	"	570
25	716	556	580
30	671	"	570
35	631	551	"
40	586	"	575
42	563	"	-
45	-	"	-
47	-	"	-
50	566	"	-
55	591	"	575
60	616	"	-
65	656	"	-
70	711	546	-
75	761	"	-
80	811	551	-
85	861	"	-
90	-	541	-
100	1000	-	720

Bakumskaya and Bergman, 1956; Bergman and Bakumskaya, 1956.

mol%	f.t.	mol%	f.t.
0	856	31	558
5	806	33	550 E
10	760	35	556
15	706	40	604
20	654	45	644
25	592	50	682
27	572 tr.t.	55	716
29	568	60	750tr.t.
		65	800
		70	842

Lithium sulfate (Li_2SO_4) + Manganese sulfate ($MnSO_4$)

Calcagni and Marotta, 1915

mol%	wt%	f.t.	E	tr.t.I	tr.t.II
0	0	856	-	-	585
3.70	5	825	570	-	-
11.40	15	755	585	-	560
19.55	25	660	"	-	565
23.80	30	610	580	-	"
28.18	35	-	"	-	"
32.70	40	590	"	-	-
37.35	45	605	"	-	565
42.15	50	620	"	-	"
47.11	55	-	585	-	570
52.22	60	630	580	-	565
57.51	65	635	595	740	560
68.62	75	650	585	810	-
74.46	80	655	575	840	-
80.51	85	660	-	845	-
86.77	90	730	-	-	-
100	100	-	-	-	-

Lithium sulfate (Li_2SO_4) + Cobalt sulfate ($CoSO_4$)

Calcagni and Marotta, 1913

%	f.t.	E	tr.t.
0	856	-	585
5	835	-	580
10	810	-	575
15	780	-	570
20	745	590	"
25	700	585	"
30	650	595	"
35	600	-	565
37	-	595	-
40	600	-	-
45	645	595	-
50	690	"	-
55	730	590	-
60	770	595	-
65	810	595	-
70	835	"	-
75	decomp.	"	-
100	-	-	-

Lesnikh and Bergman, 1953

mol%	f.t.	mol%	f.t.
60	787	27.5	590 E
55	752	25	617
50	721	20	647
45	692	15	740
40	664	10	784
35	628	5	818
30	601	0	852

Sodium acid sulfate ($NaHSO_4$) + Potassium acid sulfate ($KHSO_4$)

Rogers and Ubbelohde, 1950

%	f.t.
100	207.1
46.5	125 E
0	178.3

Sodium sulfate (Na_2SO_4) + Potassium sulfate (K_2SO_4)

Le Chatelier, 1894

mol%	f.t.	mol%	f.t.
0	860	33	830
11	830	50	855
14	825	75	940
20	815	100	1045

Jänecke, 1908

mol%	f.t.	m.t.	tr.t. I	tr.t. II
0	897	-	225	-
25	855	-	-	-
45	875	-	450	440
50	890	880	470	430
60	935	920	480	436
65	940	-	484	445
75	984	-	496	425
85	1030	-	500	425
100	1074	-	587	-

Nacken, 1918

%	f.t.	tr.t.	E	min.
0.00	883	234	-	-
6.06	861	189	-	180
11.99	848	181	181	105
17.79	834	-	182	95
23.46	830	-	183	90
34.45	839	402	186	60
44.98	856	450	-	-
50.08	867	463	-	-
55.08	883	469	-	-
59.85	895	470	-	-
64.78	914	" (1+1)	-	-
69.49	932	460	-	-
74.10	949	448	-	-
77.74	960	436	-	-
78.63	969	431	431	125
83.06	989	456	"	90
87.42	1012	498	432	60
91.69	1030	524	430	45
100.00	1076	595	-	-

Akopov and Bergman, 1954

mol%	f.t.	mol%	f.t.
0	884	42	858
5	870	45	868
10	856	48	876
15	848	51	886
20	836	55	900
24	833	60	915
27	832	65	932
30	836	70	950
33	841	75	970
36	846	100	1069
39	852		

E : 25% 832°

Bakumskaya and Bergman, 1956

%	f.t.	%	f.t.
0	884	35	846
5	868	37.5	850
10	858	40	854
15	840	42.5	860
20	836	45	864
22.5	834	47.5	872
25	832	50	880
27.5	836	52.5	888
30	838	100	1074
32.5	842		

van Hoff and Barschall 1906 and Bredig, 1942

Glaserite is a mixed crystal and not a compound

Benrath and Drekopf, 1921

t	$\varkappa \cdot 10^6$ 0 mol %	10 mol %	20 mol %	30 mol %	40 mol %
266	-	0.295	-	-	-
293	-	.479	-	-	-
320	-	1.02	-	-	-
347	1.97	.78	-	-	-
374	2.83	2.19	-	-	0.776
400	3.99	3.46	1.91	-	.997
426	6.46	4.47	3.24	2.19	1.54
452	9.77	6.46	4.90	3.89	2.88
478	14.8	10.5	8.71	8.13	7.59
504	22.9	16.2	15.1	14.8	15.1
530	34.7	23.4	21.4	22.4	25.7
555	50.1	37.2	32.4	36.8	47.9
580	75.9	60.3	56.2	81.3	118
605	118	91.2	102	159	234
630	170	145	170	282	398
655	251	214	316	550	832
679	389	316	501	1000	1050
704	575	457	813	1910	1260
728	871	692	1480	3160	1480
752	1260	1070	2750	4900	1700
776	1910	1510	4680	7240	2000
799	2880	2000	8320	9770	2400
823	4170	2750	13500	17800	2880
846	5250	4370	-	-	5750
869	10000	-	-	-	-

t	κ.10^6					
	50	60	70 mol%	80	90	100
400	0.407	-	-	-	-	-
426	.759	0.204	-	-	-	-
452	1.55	.832	-	4.47	1.62	-
478	4.07	2.51	6.31	7.59	2.63	-
504	15.1	13.8	11.0	7.59	3.89	-
530	25.1	21.4	16.2	11.5	5.75	-
555	53.7	44.7	29.5	17.8	8.91	3.16
580	123	85.1	50.1	29.5	13.8	4.47
605	229	155	87.1	45.7	20.0	6.21
630	398	246	125	64.6	27.5	8.91
655	661	363	186	91.2	38.0	12.6
679	813	525	288	126	55.0	17.4
704	1020	741	398	182	75.9	24.0
728	1260	977	562	257	105	33.9
776	1480	1200	741	372	151	47.9
799	1640	1410	1000	550	224	69.2
823	1950	1660	1230	724	331	100
846	2190	1950	1480	912	447	126
869	2460	2240	2190	1200	603	174
892	3470	2570	2460	1590	813	240
915	-	3310	3240	2090	1150	339
959	-	-	9120	2750	1590	513
981	-	-	-	3630	2190	759
1003	-	-	-	5500	3020	1150
1025	-	-	-	9770	3980	1740
1046	-	-	-	-	6030	2460
1068	-	-	-	-	9120	3390

Sodium sulfate (Na_2SO_4) + Silver sulfate(Ag_2SO_4)

Nacken, 1918

%	f.t.	tr.t.	%	f.t.	tr.t.
0.00	883	234	68.69	798	245
10.35	875	224	76.69	776	305
19.60	865	218	83.66	752	338
35.42	847	201	89.73	729	364
48.46	829	186	95.18	696	391
59.39	812	177	100.00	651	412

Belyaev and Doroshenko, 1954 (fig.)

mol%	f.t.	mol%-	f.t.
100	688	25	840
75	740	0	884
50	790		

Sodium sulfate (Na_2SO_4) + Cadmium sulfate (Cd_2SO_4)

Calcagni and Marotta, 1913

%	f.t.	E	decompos. mixed cryst.	(3+1)	(1+1)	tr.t. (1+1)	(1+3)	tr.t. (1+3)	tr.t. mix.cr.
0	887	-	-	-	-	-	-	-	235
5	897	-	266	-	-	-	-	-	224
10	896	-	-	-	-	-	-	-	-
13	894	-	346	-	-	-	-	-	224
15	895	-	371	-	-	-	-	-	-
18	893	-	-	346	-	-	-	-	-
20	891	-	-	351	-	-	-	-	-
25	876	-	421	"	-	-	-	-	224
30	861	-	466	361	-	-	-	-	"
32.84	851	-	491	351	-	-	-	-	219
35	846	-	501	346	-	-	-	-	-
40	816	-	541	361	-	-	-	-	-
45	796	-	581	-	536	-	-	-	-
50	766	-	611	-	"	416	-	-	224
55	731	-	-	-	"	466	-	-	-
60	696	681	-	-	551	496	-	-	224
63	-	"	-	-	"	491	-	-	-
65	686	"	-	-	"	"	-	-	-
68	716	"	-	-	"	496	-	-	-
70	726	686	-	-	-	500	-	-	-
72	736	676	-	-	541	496	-	-	-
75	"	-	-	-	-	481	-	-	-
77	746	-	-	-	-	-	-	456	-
79	-	-	-	-	-	-	741	446	-
80	786	-	-	-	-	-	746	456	-
81	796	-	-	-	-	-	"	461	-
83	816	-	-	-	-	-	736	456	-
85	-	-	-	-	-	-	746	446	-
90	886	-	-	-	-	-	741	436	-
100	1000	-	-	-	-	-	-	-	-

Sodium sulfate (Na_2SO_4) + Magnesium sulfate ($MgSO_4$)

Le Chatelier, 1896 and 1897

mol%	f.t.	mol%	f.t.
0	875	50	700
5	870	55	730
15	830	67	800 (1+2)
30	740	70	795
35	690	75	870 E
42	655	80	925
48	675 E	100	1170

Nacken, 1907

mol%	f.t.	m.t.	E	min.	sat.t.	(3+1)	min.
100	1124	-	-	-	-	-	-
87.7	987	-	808	120	-	-	-
78.9	884	-	809	180	-	-	-
76	811	-	808	105	-	-	-
75	813	-	-	-	-	-	-
74	799	-	658	15	-	-	-
70	776	-	667	45	-	-	-
65	746	-	662	75	-	480	60
60	672	-	665	150	-	"	90
50	683	-	665	105	-	478	120
40	726	660	-	-	633	477	135
35	754	720	-	-	592	485	157
30	789	750	-	-	550	485	172
28	796	760	-	-	538	486	180
26	807	770	-	-	515	487	195
25	814	782	-	-	506	486	210
24	816	784	-	-	500	"	165
20	836	815	200	30	485	-	-
15	853	832	212	52	463	-	-
10	865	852	210	75	427	-	-
5	876	870	221	90	318	-	-
0	883	-	-	-	234	-	-

Ginsberg, 1909

mol%	f.t.	m.t.	E	min.	(1+1) III	min.
0	883.5	883.5	-	-	-	-
2.35	879.4	-	-	-	-	-
2.9	879.0	-	-	-	-	-
5.8	875.3	863	393	-	-	-
9.3	860	-	439	-	-	-
11.6	858.7	840	473	-	-	-
15.0	844.5	-	500	-	-	-
16.2	832	-	-	-	-	-
18.2	829	800	522	-	-	-
22.8	805	-	540	-	-	-
25.1	792	755	559	-	-	-
28.3	772	732	594	-	-	-
31.3	755	-	609	-	-	-
32.5	749	-	625	-	-	-
33.7	742	-	631	-	-	-
34.75	734	-	644	-	-	-
35.8	725	670	670	-	604	-
36.4	-	-	665	2.5	"	-
39.0	700	700	674.5	-	"	-
40.0	697	697	670	-	"	-
42.1	678	678	"	6.8	603	2.0
44.2	670	670	670	14.3	603.5	3.0
45.7	678	678	"	12.0	604	5.0
47.1	-	-	"	-	599	-
47.5	700	700	674	7.0	"	-
50.0	704	704	670	8.0	"	5.3
51.2	713	713	"	5.6	604	-
52.5	729	729	"	-	"	1.3
55.0	746	746	674	-	-	-
55.2	"	"	670	5.2	607.5	1.0
56.0	747	747	"	-	603.5	-
57.0	750	750	"	-	599	-
58.4	756	756	"	-	-	-
64.0	780	780	"	3.5	599	-
67.5	806	806	"	-	-	-
71.3	817	817	661	-	599	-
75.0	850	850	814	10.9	-	-
77.0	866	866	"	-	-	-
81.2	921	921	"	8.0	-	-
88.5	995	995	"	6.0	-	-
100.0	1120	1120	-	-	-	-

mol%	(1+1) II or(3+1)	min.	(1+1) I or tr.t.1	min.	(1+1) I or tr.t.2
0	-	-	-	-	232
2.35	-	-	-	-	-
2.9	-	-	-	-	204
5.8	-	-	-	-	227
9.3	-	-	-	-	212
18.2	509	10.0	-	-	"
22.8	"	9.0	-	-	-
25.1	"	10.0	-	-	212
28.3	504.5	6.0	-	-	-
31.3	500	5.0	-	-	212
32.5	496	-	-	-	-
33.7	495	4.8	-	-	-
34.75	500	-	-	-	223
35.8	494	4.2	-	-	212
36.4	486	5.7	-	-	"
39.0	486.4	4.6	-	-	-
40.0	486	-	-	-	-
42.1	478	3.1	-	-	-
44.2	473	-	276	-	-
45.7	478	4.0	-	-	-
47.1	504.5	-	-	-	-
47.5	482.5	2.9	281	4.2	-
50.0	"	-	276	-	-
51.2	504.5	4.8	-	-	-
52.5	500.0	2.7	276	6.0	-
55.0	482.5	-	281	-	-
55.2	478.0	2.6	"	6.2	-
56.0	504.0	-	271	5.6	-
57.0	490	-	"	-	-
58.4	"	-	281	5.6	-
64.0	491	-	268	-	-
67.5	490	-	-	-	-

Sodium sulfate (Na_2SO_4) + Calcium sulfate ($CaSO_4$)

Le Chatelier, 1896 and 1897

mol%	f.t.	mol%	f.t.
0	875	40	923
1	884	46	912
3	900	51	905
5	912	54	925
7.5	923	57	950
10	930	67	1040
20	941	75	1130
30	938	100	1350

One complex.

Calcagni and Mancini, 1910

%	f.t.	E	%	f.t.
100	1375	-	29	946
80.6	1155	914	28	947
80	1152	913	27	948
71.5	1122	917	26	"
67	1080	"	24.2	949
58.4	1005	"	23	"
54.3	957	"	21	948
52.2	940	"	20	949
50.2	925	"	17.4	"
49	-	"	16.5	946
48	918	"	14.8	944
47	920	"	13.1	943
42.2	925	"	10	935
38.3	931	-	8	929
34.5	938	-	3.4	923
32.5	939	-	0.8	893
30.8	944	-	0	887

Müller, 1910

%	f.t.	m.t.	tr.t.	E
0	881	881	231	-
0.76	888	881	213	-
2.29	898	884	196	-
3.83	908	893	178	-
4.61	918	910	196	-
5.39	917	911	197	-
7.76	920	915	213	-
11.79	925	918	233	-
15.92	940	937	-	-
19.32	944	944	-	-
21.76	"	"	-	-
24.2	"	"	-	-
29.1	940	936	-	-
33.55	936	932	-	-
36.99	923	921	-	-
38.97	913	909	-	-
42.94	"	911	-	-
43.1	"	908	-	-
43.93	915	910	-	-
46.93	909	905	-	-
48.94	"	"	-	-
49.92	913	910	-	-
53.19	930	-	-	902
58.96	1018	-	-	890
63.86	1038	-	-	900
69.09	1073	-	-	880
79.3	1156	-	-	850

Bredig, 1942

No complex but mixed crystals.

Sodium sulfate (Na_2SO_4) + Strontium sulfate($SrSO_4$)

Calcagni, 1912

%	f.t.	E	mixed cryst. sat.t.	tr.t.
0	887	-	-	234
3	913	-	-	219
5	920	-	-	224
7	928	-	-	214
10	938	-	-	-
15	952	-	-	219
18	955	-	-	224
20	960	-	-	224
22	962	-	575	-
24.52	965	-	-	224
25	"	-	-	229
27	967	-	680	209
30	973	-	695	219
32	965	-	760	224
35	963	-	800	"
38	958	-	842	229
40	957	-	856	219
42	956	-	885	-
45	-	955	-	224
47	973	"	-	"
48	988	"	-	214
50	1000	"	-	"
55	1047	965	-	214
60	1075	955	-	"
65	1127	"	-	216
70	1140	"	-	"
75	1155	965	-	224
80	1300		-	-
100	1225	-	-	-

Sodium sulfate (Na_2SO_4) + Barium sulfate ($BaSO_4$)

Le Chatelier, 1897 (fig.)

mol%	f.t.	mol%	f.t.
0	875	30	955
10	900	40	1085
20	900	50	1210
25	895		

Calcagni, 1912

%	f.t.	m.t.	tr.t.	min.
0	887	-	250	-
5	895	337	230	337
10	905	620	-	620
15	913	665	230	665
20	917	750	"	750
11.6	921	770	"	770
25	917	820	"	820
29	-	913	"	-
30	917	"	"	-
32	925	"	220	-
34	940	"	230	-
35	950	"	"	-
40	987	"	"	-
45	1008	"	"	-
50	1037	"	225	-
55	1070	"	230	-
60	1110	915	220	-
65	1140	920	225	-
70	1162	925	-	-
80			-	-
100	(1345)		-	-

Sodium sulfate (Na_2SO_4) + Lead sulfate ($PbSO_4$)

Le Chatelier, 1897 (fig.)

mol%	f.t.	mol%	f.t.
0	875	52	710 E_2
10	875	60	775
20	855	70	875
30	830	80	950
40	755	90	1000
45	710 E_1	100	1035
50	715 (1+1)		

Calcagni and Marotta, 1912

%	f.t.	sat.t. (mixed cryst.)	tr.t.
0	887	-	235
3	892	-	225
5	895	325	-
7	900	365	-
10	905	435	-
13	907	-	-
15	905	505	-
20	900	565	-
25	897	620	-
30	890	675	851
35	878	735	-
40	870	730	856
42	862	730	-
48	845	-	-
50	840	750	846
52	825	750	-
55	805	735	851
60	778	735	846
63	755	-	846
65	-	735	846
68	757	735	851
70	770	735	-
75	810	735	846
80	845	735	846
100	1000	-	856

Gladushchenko and Bergman, 1955

mol%	f.t.
0	884
47.0	735
100	1100

Sodium sulfate (Na_2SO_4) + Cadmium sulfate ($CdSO_4$)

Le Chatelier, 1897 (fig.)

mol%	f.t.	mol%	f.t.
0	875	53	655 E
10	880	60	700
20	860	68	705 (1+2)
30	835	70	720 E
40	760	80	850
50	680		

Bergman and Bakumskaya, 1955

mol%	f.t.	mol%	f.t.
0	884	52.5	688
20	868	55	686
27.5	840	57.5	702
35	800	62.5	724
37.5	780	67.5	746
40	766	70	746
42.5	746	72.5	784
45	730	75	818
47.5	716	100	1000
50	702		

Bakumskaya and Bergman, 1956

%	f.t.	%	f.t.
0	884	52.5	688
20	868	55	686
27.5	840	57.5	702
35	800	62.5	724
37.5	780	67.5	746
40	766	70	746
42.5	746	70.5	746 (1+3)
43.5	736 (3+1)	72.5	784
45	730	75	818
47.5	716	100	1000
50	702		

N.B. See also Calcagni and Marotta, 1913 page 192

Sodium sulfate (Na_2SO_4) + Manganese sulfate ($MnSO_4$)

Calcagni and Marotta, 1915

mol %	f.t.	sat.t. (mixed cryst.)	(3+1)	tr.t.
0	887	-	-	235
2.82	880	-	-	215
4.72	875	-	-	-
9.46	865	360	-	210
14.24	855	420	-	205
19.01	830	-	420	-
23.82	815	485	420	-
24.84	810	-	420	205
28.73	790	515	410	210
33.62	765	545	420	220
38.55	725	575	420	-
43.50	690	-	-	-
48.47	665	630	425	-
50.41	655	-	-	-
53.49	-	645	-	-
55.50	650	645	-	-
58.53	665	645	-	-
63.60	685	645	-	-
68.70	705	650	-	-
73.83	715	640	-	780
74.87	715	-	-	800
79.00	715	-	-	845
84.20	710	-	-	845
89.45	685	-	-	850
100	-	-	-	-

Sodium sulfate (Na_2SO_4) + Cobalt sulfate ($CoSO_4$)

Calcagni and Marotta, 1913

%	f.t.	E	tr.t.
0	887	-	235
5	866	-	-
10	855	-	215
15	833	-	205
20	805	-	215
25	770	-	-
26.66	767	-	215
30	736	-	210
35.28	697	-	-
38	667	-	-
40	635	580	210
45	600	575	-
48	582	"	-
50	-	"	215
53	592	"	-
55	597	"	215
60	635	"	-
65	705	"	-
70	770	580	-
75	(850)	575	-
80	(900)	"	-
100	-	-	-

Potassium sulfate (K_2SO_4) + Silver sulfate (Ag_2SO_4)

Nacken, 1918

%	f.t.	tr.t.	E	min.
0.00	1076	595	-	-
16.58	1032	524	-	-
30.90	976	480	-	-
43.39	912	451	-	-
54.39	861	422	268	225
64.14	812	379	266	225
72.85	753	-	264	210
80.67	700	264	264	315
84.29	676	-	295	255
89.08	654	318	292	165
94.15	642	362	294	120
100.00	651	412	-	-

Potassium sulfate(K_2SO_4) + Beryllium sulfate ($BeSO_4$)

Grahmann, 1913

mol%	f.t.	E	min.	tr.t.	min.
0	1071	-	-	588	140
10	1005	758	60	"	130
20	917	757	120	586	110
30	800	764	190	588	75
40	814	768	130	585	45
50	878	765	85	581	25
60	906	758	25	565	-

E : 33 mol% (1+2)

Potassium sulfate (K_2SO_4) + Magnesium sulfate ($MgSO_4$)

Nacken, 1907

mol%	f.t.	E	min.	tr.t.	min.
100	1124	-	-	-	-
86.2	974	878	157	-	-
75.32	904	884	180	-	-
71.2	920	"	75	-	-
66.7	930	-	-	-	-
63.93	927	-	-	-	-
60	899	750	75	580	15
50	831	747	210	573	30
40	-	746	337	574	52
30	850	724	285	578	75
20	963	718	165	581	82
10	1035	-	-	570	120
0	1076	-	-	595	-

Potassium sulfate (K_2SO_4) + Magnesium sulfate (Mg_2SO_4)

Benrath and Drekopf, 1921

t	$\varkappa \cdot 10^6$						
	0	10	20	30	40	50	60
	mol%						
530	-	0.776	0.912	1.38	0.832	-	0.355
555	3.16	1.51	1.86	2.14	1.86	1.26	0.891
580	4.47	19.5	21.9	20.9	13.5	7.94	4.79
605	6.31	47.8	50	30.8	25.1	10.5	7.19
630	8.91	81.3	100	70.8	41.7	21.4	12.0
655	12.6	123	182	129	67.6	35.5	18.6
679	17.4	186	331	224	112	60.3	28.8
704	24.0	295	525	407	200	100	44.7
728	33.9	537	1000	676	363	159	69.2
	0	66.5	70	80	90	100	
728	-	0.447	-	-	-	-	
752	47.9	2.14	0.126	-	-	-	
776	69.2	6.92	.159	-	-	-	
799	140	9.55	.209	-	-	-	
823	126	12.0	.275	-	0.708	-	
846	174	15.1	.501	0.708	0.912	1.05	
869	240	19.5	.724	2.405	2.34	1.45	
892	339	28.2	-	-	-	1.86	
915	513	55.0	-	-	-	2.57	
947	759	-	-	-	-	3.55	
959	1150	-	-	-	-	4.90	
981	1740	-	-	-	-	6.61	

Potassium sulfate (K_2SO_4) + Calcium sulfate ($CaSO_4$)

Müller, 1910 and Grahmann, 1913

%	f.t.	E	min.	tr.t.	min.	(1+2)	min.	tr.t. (1+2)	min.
0	1057	-	-	580	-	-	-	-	-
3.95	1046	-	-	552	225	-	-	-	-
7.98	1038	-	-	553	210	-	-	-	-
12.11	1027	710	-	"	172	-	-	-	-
16.33	1008	-	-	555	157	-	-	-	-
20.65	982	829	45	553	150	-	-	-	-
25.08	949	837	52	551	135	-	-	-	-
29.60	914	852	105	552	105	-	-	-	-
34.24	871	861	150	554	90	-	-	-	-
36.12	867	867	165	"	"	-	-	-	-
38.99	873	865	105	"	75	-	-	-	-
43.85	918	853	45	553	60	-	-	-	-
48.84	936	848	15	550	37	-	-	-	-
50.87	947	847	"	553	30	-	-	936	7
53.95	988	845	8	552	22	-	-	935	15
57.08	994	844	8	"	15	-	-	938	22
60.97	1004	-	-	-	-	1004	67	936	34
64.57	1049	-	-	-	-	1002	30	935	22
70.09	1114	-	-	-	-	1001	22	"	15
75.75	1158	-	-	-	-	"	"	936	12

Potassium sulfate(K_2SO_4) + Strontium sulfate ($SrSO_4$)

Calcagni, 1912

%	f.t.	E	(1+1)	tr.t. (1+1)
0	1066	-	-	590
3	1074	-	-	570
5	1078	-	-	555
7	1074	-	-	565
10	1068	642	-	"
15	1060	725	-	"
20	1047	825	-	"
25	1035	970	770	"
27	1025	"	755	"
30	1013	"	770	"
32	1000	"	-	570
35	990	"	775	-
38	973	"	-	565
40	-	"	775	"
42	986	"	750	568
45	1030	"	775	565
48	1045	975	780	570
50	1070	970	770	575
55	1110	965	"	570
60	1145	960	775	565
65	1160	970	"	575
70	1180	975	780	570
75	1185	-	785	565
80	1193	-	780	"
100	1220	-	-	-

Grahmann, 1913

mol%	f.t.	m.t.	E	min.	tr.t.I	min.
0	1071	-	-	-	588	280
5	1080	1080	-	-	562	260
10	1074	1068	-	-	"	230
15	1060	1045	-	-	560	240
21.7	1039	-	-	-	558	230
30	1008	-	968	70	557	200
35	980	-	971	140	560	180
37	970	-	970	180	"	170
39	975	-	968	160	-	-
40	978	-	969	155	559	160
42	986	-	971	120	-	-
45	1023	-	970	110	559	120
48	(1060)	-	968	95	558	"
50	(1076)	-	969	75	560	105
55	1120	-	961	50	559	70
60	-	-	"	20	558	40
100	1605	-	-	-	-	-

mol%	(1+2)	min.	tr.t. II	min.
30	-	-	775	70
35	-	-	704	105
37	-	-	732	120
39	-	-	-	-
40	-	-	712	140
42	978	-	-	-
45	"	15	765	165
48	"	18	716	180
50	980	20	739	190
55	973	25	756	200
60	978	25	767	250

Potassium sulfate (K_2SO_4) + Barium sulfate ($BaSO_4$)

Calcagni, 1912

%	f.t.	E	sat.t.	tr.t.
0	1066	-	-	590
5	1075	-	-	565
10	1080	-	695	"
15	1075	-	787	"
20	1068	-	850	"
25	1055	-	935	"
30	1045	-	1005	570
35	1032	1015	-	565
40	-	1015	-	"
45	1040	"	-	570
50	1070	"	-	565
55	1090	1020	-	"
60	1125	"	-	"
65	1150	1015	-	560
70	1175	1020	-	-
100	(1345)	-	-	-

Grahmann, 1913

mol%	f.t.	m.t.	E	min..	tr.t..	min..	sat. t.
0	1071	-	-	-	588	280	-
10	1081	1081	-	-	572	210	783
20	1064	1040	-	-	"	170	961
30	1014	-	996	180	576	120	-
40	1060	-	1016	150	568	100	-
50	1167	-	"	120	571	"	-
60	1252	-	"	110	572	80	-
100	1580	-	-	-	-	-	

Potassium sulfate (K_2SO_4)+ Lead sulfate ($PbSO_4$)

Calcagni and Marotta, 1912

%	f.t.	E	(1+1)	decompos. mixed cryst.	tr.t. I	tr.t. II
0	1066	-	-	-	590	-
5	1063	-	-	565	550	-
10	1058	-	-	-	"	-
15	1046	-	-	560	"	-
20	1028	-	-	580	"	-
25	1013	-	-	600	"	-
30	1003	-	610	670	560	-
35	981	-	615	760	540	-
40	956	791	"	-	550	-
45	921	789	"	-	540	-
50	876	796	620	-	550	866
55	836	"	"	-	-	-
60	-	"	615	-	550	856
63.48	834	"	620	-	560	-
65	856	"	"	-	550	856
70	906	-	615	-	"	-
75	933	801	620	-	540	856
80	949	-	625	-	550	840
100	1010	-	-	-	-	856

Grahmann, 1913

mol%	f.t.	E	min.	tr.t. I	min.	(1+1)	min.
0	1071	-	-	588	140	-	-
10	1044	-	-	538	"	-	-
20	1005	-	-	537	110	593	-
30	946	780	-	"	60	619	35
40	851	"	70	535	30	612	130
50	842	792	100	537	-	617	180
60	916	"	25	-	-	603	65
70	935	837	15	-	-	-	-
80	849	836	70	-	-	-	-
90	975	"	50	851	25	-	-
100	-	-	-	852	60	-	-

mol%	tr.t. II	min.	mol%	tr.t. II	min.
50	537	-	70	544	160
60	536	160	80	536	130

Potassium sulfate (K_2SO_4) + Cadmium sulfate ($CdSO_4$)

Calcagni and Marotta, 1913

%	f.t.	E	(1+2)	tr.t. I	tr.t. II
0	1066	-	-	-	590
5	1053	-	-	553	513
10	1034	-	-	-	528
15	1013	-	-	-	"
20	983	-	-	-	523
25	945	-	-	-	"
28.5	908	-	-	-	526
30	903	-	-	598	528
35	848	-	-	628	523
40	788	640	-	-	528
45	713	653	-	-	508
48		"	-	-	513
50	-	"	-	-	523
52	663	-	-	-	533
55	673	658	-	-	523
58	693	653	-	-	"
60	713	643	-	-	533
65	753	648	-	-	-
68	763	-	-	-	523
70	788	-	763	-	-
72	788	-	"	-	-
75	803	-	753	-	-
78	813	-	(1+3)	-	-
80	823	-	813	-	-
82	833	-	"	-	-
85	863	-	"	-	-
90	933	-	788	-	-
95	973	-	"	-	-
100	1000	-	-	-	820

Bakumskaya and Bergman 1956;
Bergman and Bakumskaya, 1956.

%	f.t.	%	f.t.
0	1074	55	746
25	912	58	758
27.5	888	60.5	764 (1+2)
29.5	856 tr.t.	61	766
30	856	63	776
32.5	836	67	776
35	810	70	806
40	728	75	818 (1+3)
42.5	686	80	822
43.5	653 E	85	900
45	680	100	1000
50	720		

Potassium sulfate(K_2SO_4) + Manganese sulfate ($MnSO_4$)

Calcagni and Marotta, 1915

mol%	wt%	f.t.	E	tr.t.
0	0	1066	-	590
5.73	5	1045	-	560
11.37	10	1010	585	"
16.92	15	975	630	565
22.38	20	925	675	560
27.78	25	855	680	"
33.09	30	775	"	"
38.32	35	710	"	570
40.40	37	-	"	560
43.48	40	700	-	-
48.56	45	750	685	560
52.58	49	775	680	-
53.54	50	785	685	-
58.51	55	815	680	-
63.38	60	835	675	-
68.18	65	845	-	-
72.92	70	830	-	-
74.79	72	825	-	-
77.58	75	815	-	-
80.36	78	805	-	-
82.19	80	800	-	-
86.74	85	815	-	-
100	100	-	-	-

Potassium sulfate (K_2SO_4) + Cobalt sulfate ($CoSO_4$)

Calcagni and Marotta, 1913

%	f.t.	E	tr.t.	
0	1066	-	590	
5	1045	515	560	
10	1010	535	"	
15	965	"	570	
20	920	530	565	
25	848	535	560	
30	740	540	"	
35	640	535	"	
38	590	-	-	
40	-	535	-	
42	550	"	-	
45	560	-	-	
48	635	-	-	
50	655	-	-	
55	700	-	-	
60	720	-	-	(1+1)
62	723	-	-	
64	736	-	-	
65	732	-	-	
68	725	725	-	
70	740	"	-	
72	775	720	-	
75	835	725	-	
80	-	"	-	

Rubidium sulfate (Rb_2SO_4) + Magnesium sulfate ($MgSO_4$)

Plyushchev and Markovskaya, 1954

mol%	f.t.	E	tr.t.	(1+2)
0	1074	-	645	-
1.0	1058	676	"	-
2.5	1040	675	646	-
5.0	1026	677	643	-
10.0	995	672	645	-
15.0	955	675	"	-
20.0	918	"	648	-
25.0	882	672	645	-
30.0	830	678	"	-
35.0	775	675	648	-
40.0	728	"	"	-
45.0	700	673	645	-
46.5	685	675	"	-
47.5	675	"	"	-
50.0	728	"	648	-
52.5	785	"	"	-
55	800	"	642	-
57.5	830	678	645	-
59.0	850	675	"	850
60.0	860	"	"	"
62.5	870	"	"	"
65.0	880	-	-	"
66.7	890	--	-	"
70.0	898	-	-	853
75.0	910	-	-	847
80.0	930	-	-	850
85.0	975	-	-	848
90.0	1015	-	-	846
95.0	1065	-	-	850
100.0	1124	-	-	-

Rubidium sulfate (Rb_2SO_4) + Calcium sulfate ($CaSO_4$)

Müller, 1910

%	f.t.	E	min.	tr.t.	min.
0	1051	-	-	649	-
2.61	1039	-	-	631	105
5.36	1032	-	-	625	90
11.31	971	824	45	626	75
14.53	942	827	60	624	52
17.93	902	833	90	625	37
20.31	870	840	127	629	30
23.43	840	"	157	"	"
25.37	846	846	"	627	22
29.44	881	845	75	622	"
33.77	922	"	60	626	15
38.39	949	"	37	621	"
43.34	973	829	22	623	"
46.47	1019	837	"	628	8
50.49	1043	-	-	-	-
54.33	1058	-	-	-	-
60.47	1104	1055	-	-	-
67.10	1156	1038	-	-	-

%	transitions of complex			
	I	min.	II	min.
38.39	909	22	-	-
46.47	911	30	781	-
50.49	911	30	788	30
54.33	915	30	787	45
60.47	910	27	797	41
67.10	896	24	852	36
67.10	866	70	866	70

Cesium sulfate (Cs_2SO_4) + Magnesium sulfate ($MgSO_4$)

Plyushchev and Markovskaya, 1954

mol%	f.t.	E	tr.t.	(1+1)	(2+1)
0.0	1019	-	660	-	1019
1.0	1008	680	660	-	1008
2.5	1000	682	662	-	1000
5.0	995	680	660	-	995
10.0	968	"	663	-	968
15.0	940	"	662	-	940
20.0	920	"	660	-	920
25.0	895	"	661	-	895
30.0	843	"	660	-	843
33.3	820	"	"	-	820
35.0	800	"	"	-	800
40.0	740	676	661	-	570
43.0	700	678	660	-	"
45.0	680	680	661	-	"
47.0	696	"	"	-	"
50.0	726	"	660	-	"
55.0	775	"	657	-	"
60.0	810	676	660	-	566
63.0	825	680	658	825	570
65.0	840	682	660	"	"
70.0	890	680	655	"	"
75.0	920	-	-	"	"
76.0	932	-	-	"	"
80.0	965	-	-	826	"
85.0	1022	-	-	825	"
90.0	1050	-	-	826	"
95.0	1085	-	-	825	"
100.0	1124	-	-	-	"

Thallium sulfate (Tl_2SO_4)+ Cadmium sulfate ($CdSO_4$)

Sementsova, Bergman and Lesnykh, 1956.

mol%	f.t.	mol%	f.t.
0	638	31	514 E
5	626	35	562
10	606	40	598
15	588	45	642
20	564	50	680
25	544	55	726
27.5	528	60	764
29	518	65	800
30	514		

Sodium chromate (Na_2CrO_4) + Potassium chromate (K_2CrO_4)

Amadori, 1913

mol%	f.t.	tr.t.	E
0	780	392	-
2.5	775	380	-
5	764	-	366
10	757	-	"
15	748	-	364
20	745	494	363
25	740	522	360
30	743	"	"
40	759	592	-
50	780	607	-
60	818	604	-
70	859	592	-
75	878	575	-
77.5	892	570	-
80	902	-	550
82.5	914	-	552
85	922	570	"
90	938	607	550
95	956	637	-
100	978	666	-

Sodium chromate (Na_2CrO_4) + Calcium chromate ($CaCrO_4$)

Vilnyanski and Pudovkina, 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	795	50	750
15	812	70	910
30	790	78	975
48.4	740 E		

Lithium molybdate (Li_2MoO_4) + Sodium molybdate (Na_2MoO_4)

Hoermann, 1928

%	f.t.	E	min.	(1+3)	min.	tr.t.	min.
0	705	-	-	-	-	-	-
10	663	445	35	-	-	-	-
20	621	456	65	-	-	-	-
30	676	460	108	-	-	-	-
40	820	464	275	-	-	-	-
48	473	466	220	-	-	-	-
50	466	464	208	-	-	-	-
52	-	"	190	-	-	-	-
55	478	"	155	-	-	-	-
60	489	455	100	484	18	-	-
63	502	439	60	"	20	-	-
65	511	430	48	483	60	-	-
67	519	421	25	"	75	-	-
70	532	-	-	481	90	-	-
73	545	-	-	482	92	-	-
75	555	-	-	481	102	415	25
80	580	-	-	467	70	421	85
90	630	-	-	467	30	423	210
100	686	-	-	-	-	"	360

Lithium molybdate (Li_2MoO_4) + Potassium molybdate (K_2MoO_4)

Hoermann, 1928

%	f.t.	E	min.	tr.t.	min.
0	705	-	-	-	-
10	665	522	60	398	25
20	610	521	132	406	60
30	544	522	200	409	75
32.5	522	"	264	408	84
35	534	521	192	"	110
40	554	"	95	410	115
45	566	"	50	"	120
50	571	-	-	412	122
55	564	548	90	406	102
60	550	550	187	398	90
70	678	"	145	398	65
75	733	549	120	395	40
80	784	547	90	400	25
90	857	543	35	-	-
100	926	-	-	-	-

Bergman, Kislova and Korobka, 1954

mol%	f.t.	mol%	f.t.
0	705	40	546
10	658	45	561
12.5	641	50	569 (1+1)
14.5	630 tr.t.	55	559
15	628	59.5	547 E
18	611	60	559
20	602	64	619
25	578	70	691
30	543	80	783
33.5	513 E	90	859
35	523	100	924

Lithium molybdate (Li_2MoO_4) + Silver molybdate (Ag_2MoO_4)

Belyaev and Doroshenko, 1956

mol %	f.t.	mol %	f.t.	mol %	f.t.
100	554	67.5	448	35	576
95	540	65	460	30	592
90	527	62.5	473	25	609
85	512	60	485	20	628
80	496	55	504	15	644
75	480	50	524	10	662
72.5	467	45	540	5	679
70	456	40	558	0	700
68	445 E				

Lithium molybdate (Li_2MoO_4) + Cadmium molybdate ($CdMoO_4$)

Lesnykh, Bergman and Bukun, 1956 (fig.)

mol %	f.t.	mol %	f.t.
0	700	25	700
10	690	35	750
19.5	654 E	40	840

Sodium molybdate (Na_2MoO_4) + Potassium molybdate (K_2MoO_4)

Amadori, 1913

%	f.t.	tr.t.
0	688	614-574-405
2.5	682	550-250
5	676	485
7.5	674	459
10	670	430
15	664	400
20	660	378
25	667	358
30	672	340
40	692	304
50	724	682
60	763	667
70	803	644
75	820	628
80	844	598
85	863	570
87.5	875	547
90	886	528
95	905	460
98	918	460
100	926	475

Belyaev and Sholokhovich, 1953 (fig.)

mol%	f.t.	mol%	f.t.
100	926	21	656 E
75	820	0	690
50	720		

Bukhalova and Mateiko, 1955 and Mateiko and Bukhalova, 1955

mol%	f.t.	mol%	f.t.
0	688	40	697
2	"	46	715
4	"	50	727
6	687	54	736
8	686	55	743
10	684	58	760
13	680	62	775
16	674	66	788
19	667	70	806
22	672	74	821
25	675	78	834
28	677	82	856
31	678	86	874
34	684	90	884
37	688	100	926

E : 19 mol% 667°

Sodium molybdate (Na_2MoO_4) + Silver molybdate (Ag_2MoO_4)

Belyaev and Doroshenko, 1954

mol%	f.t.	mol%	f.t.
100	554	55	597
95	553	50	611
90	550	45	622
85	548	40	632
82.5	547	35	640
80	546	30	648
77.5	"	25	656
75	548	20	662
72.5	550	15	668
70	553	10	674
67.5	560	5	681
65	566	0	688
60	580		

Potassium molybdate (K_2MoO_4) + Lead molybdate ($PbMoO_4$)

Belyaev, Sholokhovich and Barkova, 1956.

mol%	f.t.	mol%	f.t.
0	926	50	870 (1+1)
15	850 tr.t.	63	802
31	770 E	100	1055

Lead molybdate ($PbMoO_4$) + Bismuth molybdate ($Bi_2Mo_3O_{12}$)

Zambonini, 1920

mol%	f.t.	E	min.
100	632	-	-
95	638	610	50
90	634	646	80
80	628	"	150
70	627	615	180
60	734	614	170
50	820	"	150
40	872	608	110
30	910	610	90
20	977	"	70
10	1022	608	40
5	1054	600	20
0	1065		

Lead molybdate ($PbMoO_4$) + Didymium molybdate ($Di_2Mo_3O_{12}$)

Zambonini, 1916

%	f.t.	m.t.	%	f.t.	m.t.
0	1065	-	50	1094	1080
5	1069	1065	60	1106	1085
10	1072	"	70	1115	1090
20	1076	-	80	1123	1100
30	1079	1070	90	1137	-
40	1084	1075	100	1144	1125

Lead molybdate ($PbMoO_4$) + Praseodymium molybdate ($Pr_2Mo_3O_{12}$)

Zambonini, 1915

%	f.t.	m.t.	%	f.t.	m.t.
0	1065	-	60	1054	1050
10	1072	1068	70	1048	1046
20	"	"	80	1043	1040
30	1070	1067	90	1038	1035
40	1068	1064	100	1030	-
50	1061	1056			

Mixed crystals with a maximum f.t.

Lead molybdate ($PbMoO_4$) + Neodymium molybdate ($Nd_2Mo_3O_{12}$)

Zambonini, 1916

%	f.t.	m.t.	%	f.t.	m.t.
0	1065	-	60	1115	1105
10	1071	1068	70	1131	1120
20	1076	1070	80	1147	1135
30	1080	1075	90	1162	1150
40	1088	1080	100	1176	-
50	1102	1090			

Lead molybdate ($PbMoO_4$) + Cerium molybdate ($Ce_2Mo_3O_{12}$)

Zambonini, 1915 and 1916

%	f.t.	m.t.	E
0	1065	-	-
5	1067	-	-
10	1071	-	-
13.8	1072	-	-
20	1074	-	max.
30	1064	-	-
40	1055	1035	-
50	1054	1025	-
60	1034	1010	-
70	1024	1000	-
80	1004	980	-
90	975	-	938
95	948	-	"
100	978	-	-

Lead molybdate ($PbMoO_4$) + Lanthanum molybdate ($La_2Mo_3O_{12}$)

Zambonini, 1915 and 1916

%	f.t.	m.t.	%	f.t.	m.t.
0	1065	-	60	1139	1134
10	1074	1067	70	1151	1145
20	1085	1078	80	1163	1156
30	1095	1087	90	1175	1169
40	1109	1102	100	1181	-
50	1126	1120			

Lead molybdate ($PbMoO_4$) + Yttrium molybdate ($Y_2Mo_3O_{12}$)

Zambonini, 1915 and 1916

%	f.t.	E	min.
100	1347	-	-
95	1324	-	-
90	1296	974	20
80	1235	984	100
70	1143	974	150
60	1038	984	200
50	1000	976	180
40	1016	973	160
30	1030	977	100
20	1038	"	70
10	1046	970	30
0	1065	-	-

Lithium tungstate (Li_2WO_4) + Silver tungstate (Ag_2WO_4)

Lesnykh and Bergman, 1956

mol%	f.t.
0	736
55	560 E
100	584

Lithium tungstate (Li_2WO_4) + Sodium tungstate (Na_2WO_4)

van Liempt, 1925

mol%	f.t.	E	mol%	f.t.	E
0	742	-	50	504	489
5	725	489	60	537	487
10	711	"	70	577	"
20	667	487	80	622	"
30	590	489	85	645	490
40	551	"	90	660	-
45	501	491	100	705	-

Hoermann, 1929

%	f.t.	E	min.	(1+3)	min.	tr.t.	min.
0	742	-	-	-	-	-	-
10	710	488	36	-	-	-	-
20	665	483	95	-	-	-	-
30	615	488	180	-	-	-	-
40	552	490	220	-	-	-	-
45	514	492	275	-	-	-	-
50	495	492	300	-	-	-	-
55	505	491	235	-	-	-	-
60	520	475	190	511	15	-	-
70	564	472	65	"	60	-	-
75	589	471	25	510	85	-	-
80	612	-	-	511	60	574	35
90	658	-	-	510	30	579	90
100	700	-	-	-	-	583	178

Lithium tungstate (Li_2WO_4) + Potassium tungstate (K_2WO_4)

van Liempt, 1925

mol%	f.t.	E	mol%	f.t.	E
0	739	-	55	627	-
10	697	562	60	607	607
15	667	572	65	-	"
20	642	582	70	689	602
25	602	572	75	732	607
30	579	-	80	782	592
35	602	572	85	817	600
40	617	"	90	854	605
45	627	"	100	933	-
50	632	-			

Bergman, Kislova and Posipaiko, 1954 and 1955

mol%	f.t.	mol%	f.t.
0	738	55	618
5	716	57.5	610
10	690	60	600
12.5	678	62.5	582
15	656	65	588
17.5	648	67.5	624
20	644	70	654
22.5	632	72.5	682
25	618	75	712
27.5	598	77.5	740
30	586	80	760
32.5	566	82.5	782
35	574	85	800
37.5	586	87.5	826
40	600	90	846
42.5	610	92.5	864
45	618	95	890
47.5	622	97.5	910
50	626	100	926
52.5	622		

Lithium tungstate (Li_2WO_4) + Lead tungstate ($PbWO_4$)

Belyaev, 1955

mol%	f.t.	mol%	f.t.
0	738	35	816
5	729	40	841
10	720	45	871
15	710	50	891
20	729	55	925
25	760	60	950
30	786	65	975

Sodium tungstate (Na_2WO_4) + Potassium tungstate (K_2WO_4)

Amadori, 1913

mol%	f.t.	tr.t. I	tr.t. II	E
0	674	576-560	-	- •
5	676	558-546	-	-
10	658	548-525	-	-
15	647	545-512	-	-
20	636	543-474	-	-
25	630	542-459	-	-
30	634	-	-	-
40	650	536-432	-	-
50	672	530-400	-	-
55	684	512-362	658	-
60	704	-	650	-
70	740	-	642	-
75	758	-	625	-
80	782	-	612	-
85	882	-	590	522
90	826	-	570	"
95	864	-	535	524
100	894	575	575	-

van Liempt, 1925

%	f.t.	tr.t. II-III	tr.t. III-IV
0	705	588	577
10	660	-	546
15	650	-	"
20	646	-	530
30	660	-	470
40	670	-	420
50	710	-	-
60	740	-	-
70	784	-	-
80	820	-	-
90	850	-	-
100	926	358	-

Sholokhovich and Bergman, 1954

mol%	f.t.	mol%	f.t.
100	927	49	705
97	915	46	696
94	910	42	689
91	895	39	684
88	876	36	675
85	865	33	671
82	850	30	665
79	840	27	661
76	827	24	651
73	815	21	645
70	793	18	645
67	783	15	645
64	770	12	651
61	756	9	655
58	742	6	680
55	727	3	689
52	718	0	694

(1+1) 689° (45 mol%) incongruent

Bukhalova and Mateiko, 1956.

mol%	f.t.		mol%	f.t.
0	698		42.5	686
5	677		45	698
10	658		50	722
12.5	654		55	744
15	645		60	770
15.5	642	E	65	790
17.5	645		70	810
20	645		75	834
22.5	645		80	853
23.5	646	tr.t.	82.5	860
25	654		85	866
30	662		87.5	882
35	670		90	894
37.5	674		97.5	916
40.4	676		100	926
40	678	tr.t.		

Lead tungstate ($PbWO_4$) + Bismuth tungstate ($Bi_2W_3O_{12}$)

Zambonini, 1920

mol%	f.t.	E	min.
100	832	-	-
90	829	810	140
80	820	813	200
70	816	"	230
60	845	814	190
50	877	813	140
40	919	814	130
30	966	"	110
20	1023	"	95
10	1075	806	70
7	1093	804	55
5	1109	796	45
0	1130	-	-

Lead tungstate ($PbWO_4$) + Cerium tungstate ($Ce_2W_3O_{12}$)

Zambonini, 1915 and 1916

%	f.t.	m.t.	%	f.t.	m.t.
100	1089	-	40	1122	1108
90	1096	1091	30	1122	1108
80	1103	1092	20	1123	1108
70	1112	-	10	1123	1110
60	1115	-	0	1125	-
50	1119	1098			

Sodium dichromate ($Na_2Cr_2O_7$) + Potassium dichromate ($K_2Cr_2O_7$)

Lehrman, Selditch and Hell, 1936

%	f.t.	E	%	f.t.
100.0	397.4	-	85	372
84.7	373	304	72	351
69.1	351	304.6	61	333
57.9	328	304	50	316
53.2	317	303	45	311
37.9	-	304.6	37	303
26.9	-	306	29	310
24.9	-	304	21	320
22.9	315	304.6		
14.6	335	-		
0	356.7	-		

Lithium orthosilicate (Li_4SiO_4)
+ Calcium orthosilicate (Ca_2SiO_4)

Schwarz, 1921

mol%	d	f.t.	m.t.
0	2.280	1249	-
10	.400	1199	1171
20	.423	1145	1091
25	.436	-	-
30	.424	1091	E
35	.478	-	-
40	.591	1107	(3+2)
45	.636	-	-
50	.647	1071	E
55	.669	-	-
60	.666	1091	(2+3)
70	-	1052	E
100	.97	-	-

Lithium orthosilicate (Li_4SiO_4)
+ Aluminum orthosilicate ($Al_4Si_3O_{12}$)

Ballo and Dittler, 1912

mol%	f.t.	m.t.	E	(3+1)	d	E
0	1215	1215	-	-	2.2800	-
10	1183	1023	1023	1072	.3627	2.3973
20	1145	1014	1014	1074	.3651	.4042
25	1080	1018	1018		.5172	.5110
30	1050	1022	1022	-	.5158	.4961
35	1120	1023	1023	-	.405	-
45	1314	-	-	-	.200	-
50	1330	1330	-	-	.200	-
60	1265	1170	-	-	.426	-
70	1285			-	.474	-

$(3+1)^{-}$ and $(1+\bar{1})$

Lithium orthosilicate (Li_4SiO_4)
+ Zirconium orthosilicate ($ZrSiO_4$)

Schwartz, 1921

mol%	f.t.	E	d	
0	1249	-	2.280	
10	1189	996	3.145	
20	1088	1000	.356	
25	1068	1006	.382	
30	-	1021	.401	
33.4	1030	1019	.410	
35	1036	1016	.449	
40	1104	1021	-	(3+2)
50	1143	1008	.479	
60	1152	-	.601	
65	1106	-	4.024	
66.7	1099	-	3.905	
70	1062	-	4.124	
100	2000	-	.51	

Sodium orthosilicate (Na_4SiO_4)
+ Potassium orthosilicate (K_4SiO_4)

Thomas, 1910

%	f.t.	%	f.t.
0	1040	60	870
10	1010	70	860
20	960	75	853 E
30	900	80	860
40	820	90	878
50	743 E	100	913
56	873		

one complex

Magnesium orthosilicate (Mg_2SiO_4)
+ Calcium orhtosilicate (Ca_2SiO_4)

Deleano, 1913

mol %	m.t.	mol %	m.t.
100	2130	40	1435
90	1655	20	1400
80	1650	10	1515
70	1470	5	1560
55	1495	0	1560
50	1465		

Olivine (Mg_2SiO_4) + Augite ($(Ca,Mg,Fe)_2SiO_4$)

Vučnik, 1904

%	f.t. glass
0	1300
50	1170
75	1160
80	1170
90	1160-1170
100	1185

%	f.t. mixed crystals
0	1280
80	1080
100	1185

Magnesium orthosilicate (Mg_2SiO_4)
+ Manganese orthosilicate (Mn_2SiO_4)

Kallenberg, 1914

mol%	f.t.	d
100	1210	4.02
90	1240	3.92
80	1280	3.84
70	1330	3.79
60	1380	3.70
50	1430	3.54
0	-	3.27

Calcium orthosilicate (Ca_2SiO_4)
+ Barium orthosilicate (Ba_2SiO_4)

Toropov and Konovalov, 1938

mol%	d	n_D ordin.	n_D extraordin.
	25°		
0	3.28	1.735	1.717
4.9	.51	.737	.719
16.8	.83	,720	.725
31.7	4.18	.757	.737
51.8	4.4	.763	.749
73.4	4.8	.783	.768
100	5.2	.830	.810

Calcium orthosilicate (Ca_2SiO_4)
+ Manganese orthosilicate (Mn_2SiO_4)

Kallenberg, 1914

mol%	f.t.	d
100	1210	4.09-4.02
90	1190	3.85
80	1220	.76
65	1230	.60
50	1250	.39
40	1280	.34
30	-	.23
20	-	decomposition
10	-	"
0	2130	"

Takody, 1928

mol%	f.t.	m.t.	d
100	1210	1170	4.05
90	1190	1160	3.86
80	1210	1160	.64
70	1220	1200	.56
50	1240	1220	.39
10	1550	-	.04

Calcium orthosilicate (Ca_2SiO_4)
+ Ferrous orthosilicate (Fe_2SiO_4)

Bowen, Schairer and Posnjak, 1933 (fig.)

%	f.t.	m.t.	%	f.t.	m.t.
100	1205	1204	48	1226	1215
88	1160	1125	44	1300	1220
81=	1117	1117	39	1425=	1230
77	1122	1120	34	1525	1228
66	1182	1150	15	-	1230
54.5	1208	1208 (1+1)	0	2130	2130

Sodium aluminum orthosilicate ($NaAlSiO_4$)
+ Potassium aluminum orthosilicate ($KAlSiO_4$)

Bowen, 1917

%	f.t.	m.t.	tr.t.
0	1530	-	1250
10	1485	1425	1375-1305
16	1465	1405	1330
27.5	1405	1405	1380
40	1475	1415	-
50	1525	1440	-
60	1580	1490	-
70	1640	1550	

Sodium aluminum orthosilicate ($NaAlSiO_4$)
+ Calcium aluminum orthosilicate ($CaAl_2Si_2O_8$)

Bowen, 1912

%	f.t.	E	tr.t.
0	1526	1305	-
10	1479	1335	1353
20	-	-	1356
30	-	-	1349
40	-	1310	-
50	-	1305	-
60	-	1302	-
70	1437	1307	-
80	1484	1306	-
90	1522	-	-
100	1550	-	-

%	mixed crystals σ	mixed crystals ε
0	1.537	1.533
10	"	.535
20	"	.537
23	"	.537
25	"	.537
35	"	.539

Smalley, 1947 (fig.)

%	f.t.	tr.t. 1	tr.t. 2	tr.t. 3
100	1590	-	-	-
70	1470	-	-	-
60	1420	-	-	-
50	1358	1285	1230	-
45	1325	1285	1265	-
40	1300	1285	1285	-
36.5	1285	1285	1290	1220
35	1282	1280	1290	1220
32.5	1295	1275	1295	-
30	1325	1275	1300	1195
20	1400	1270	1300	1155
10	1465	1230	1280	1150
0	1525	-	-	-

%	n	%	n
100	1.636	40	1.560
80	.610	20	.535
60	.585	0	.515

Acmite ($Fe_2Na_2O_{12}Si_4$) + Leucite ($KAlSi_2O_6$)

Vucnik, 1904

%	f.t.
0	965
16.67	880
25	890-900
100	1310

Calcium Aluminium orthosilicate ($CaAl_2Si_2O_8$) + Barium Aluminium orthosilicate ($BaAl_2Si_2O_8$)

Ginsberg, 1911

%	f.t.
0	1490
50	1323 E
100	1640

Lithium disilicate ($Li_2Si_2O_5$) + Sodium disilicate ($Na_2Si_2O_5$)

Kracek, 1939

%	f.t.	tr.t.	E
100	874	693	-
95	845	"	-
90	815	"	641
80	760	"	"
75	735	"	"
73	701	-	"
68	710	-	"
60	785	-	"
50	850	-	"
40	910	-	"
30	960	-	"
20	985	-	-
10	1020	-	-
0	1034	-	-

Albite ($NaAlSi_2O_6$) + Diopside ($CaMgSi_2O_6$)

Bowen, 1916

%	f.t.	%	f.t.
100	1390	25	1210
75	1340	10	1145
50	1285	0	1100

Leucite ($KAlSi_2O_6$) + Diopside ($CaMgSi_2O_6$)

Bowen and Schairer, 1929

%	f.t.	%	f.t.
0	1686	60	1312
10	1600	61.5	1300
20	1534	70	1319
30	1477	80	1336
40	1425	90	1363
50	1377	100	1391.5

Lithium acetate ($C_2H_3O_2Li$) + Sodium acetate ($C_2H_3O_2Na$)

Diogenov, 1956.

mol%	f.t.	mol%	f.t.
100	337	43.5	165
98	333	43	160E
96	326	41.5	169
-	323 tr.t.	38.5	185
93	320	34	206
90	316	29.7	216
86	310	24.5	225
82	303	20	227 (4+1)
77.7	295	18	232
74	285	14	248
70.7	277	9.7	253
67	266	-	257 tr.t.
58.5	237	5.5	266
54.5	222	4	277
50.8	208	3	284
47	188	0	291

Lithium acetate ($C_2H_3O_2Li$) + Potassium acetate ($C_2H_3O_2K$)

Diogenov, 1956.

mol%	f.t.	mol%	f.t.
100	310.5	47.5	247
97.5	308	44.8	258
96.3	305	41.5	268
92.5	295	37.5	273
88.5	285	32.5	275 (2+1)
84.5	274	32.2	274
80.7	262	27.5	267
77.5	251	23.7	257
73.3	234	20.0	245
70.5	221	17.7	235
68.3	209	16.0	225
66.0	197	15.5	222 E
64.5	187	15.0	229
64	181 E	13.3	240
62.5	191	10.6	252
60.0	208	7.0	262
57.5	221	5.5	270
54.0	232	4.0	278
50.0	236 (1+1)	0.0	291
48.5	236 E		

Sodium acetate ($C_2H_3O_2Na$) + Potassium acetate ($C_2H_3O_2K$)

Baskow, 1915

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	295.0	-	48	-	232.0
88	288.0	263.5	41.5	253.0	240.5
76	260.0	250.0	33.5	271.5	256.5
62.5	237.0	228.5	22	293.0	277.2
59.5	231.0	223.5	12.5	307.5	295.0
53.5	223.0	223.0	0	320.0	-

De Mol, 1925 (fig.)

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	306	306	50	248	250
90	290.8	297.0	40	258.2	280.2
80	276.2	288	30	273.2	306
70	262.5	276.5	20	290	317.4
60	251	262	10	308.4	324.3
53	248.5	248.5	0	330	330

Potassium acetate ($C_2H_3O_2K$) + Cadmium acetate ($CdC_4H_6O_4$)

Lehrman and Schweitzer, 1954

mol%	f.t.	E	mol%	f.t.	E
0.0	292	-	42.86	221	-
10.0	289	-	44.44	216	201
20.0	246	183	48.0	206	-
30.0	195	-	50.0	210	-
33.3	202	-	52.0	205	-
35.0	196	188	55.0	202	187
38.0	203	-	60.0	190	-
40.0	213	-	70.0	220	-
41.0	217	-			

(2+1) (1+1) (4+3)

Sodium rubidium tartrate ($NaRbC_4H_4O_6$) + Sodium ammonium tartrate ($NaNC_4H_8O_6$)

Kozik, 1931

w.l.	n_D 17° α	β crystal axes	
	87%		
670.8	1.4887	1.4948	1.4958
656	.4895	.4956	.4964
610	.4907	.4973	.4978
589	.4915	.4983	.4992
546	.4942	.5007	.5012
527	.4945	.5012	.5015
486	.4977	.5055	.5057
470	.4993		
436	.5028	1.5106	1.5107
	77%		
670.8	1.4891	1.4947	1.4955
656	.4896	.4953	.4958
610	.4910	.4972	.4977
589	.4920	.4983	.4986
546	.4943	.5007	.5007
540	.4946		
527	.4953	1.5020	1.5022
486	.4984	.5052	.5054
436	1.5034	.5102	.6107

Magnesium nitrate h. (MgN_2O_6). 6 H_2O
+ Lanthanum nitrate h. (LaN_3O_9).6 H_2O

Quill and Robby, 1937

%	f.t.	%	f.t.
0.0	89.7	60.0	110.8
5.0	88.5	67.2	110.6
E	88.0	70.0	107.8
10.0	91.0	78.0	101.6
20.0	101.7	90.0	76.5
30.0	107.7	95.0	64.7
40.0	111.5	100	66.5
50.0	113.2		

(3+2) h= hydrate

Calcium nitrate h. (CaN_2O_6).4 H_2O
+ Cadmium nitrate h. (CdN_2O_6).4 H_2O

Hasselblatt, 1913

%	f.t.	m.t.	
100	59.3	-	Mixed crystals
78.7	54.1	48.1	"
61.8	49.2	42.8	"
51.8	46.3	41.1	"
19.8	42.4	39.7	"
13.2	40.9	39.6	"
0	39.7	-	"

%	f.t.
20.2	37.8
13.2	39.8
3	41.7
0	42.6

%	S.C.*	%	S.C.*
	Mixed crystals		
100	30.4	20.2	12.8
88.7	21.7	13.2	12.3
51.8	14.5	0	12.2
29.8	13.7		
	Calcium nitrate		
29.8	1.9	3	11.3
20.2	4.0	0	13.6
13.2	6.2		

* speed of crystallization

Hasselblatt, 1921

P	f.t.	P	f.t.
	0%		
1	42.5	1590	57.55
1196	54.8	3050	68.95

P	f.t.	m.t.	P	f.t.	m.t.
		5.3 %			
1	40.2	39.6	510	42.6	-
		11.4 %			
1	41.0	39.6	1240	47.2	-
1150	-	45.4	2055	49.9	-
		29.1 %			
1	43.2	40.0	2290	57.4	-
840	-	46.7	2830	-	58.7
970	49.45	-	3075	62.0	-
2070	-	53.8			
		53.7 %			
1	47.0	41.5	2085	-	56.95
300	-	46.0	2210	61.75	-
440	51.65	-	3000	-	60.4
900	-	51.2	3095	65.25	-
1010	59.45	-			

P	f.t.	P	f.t.
	100%		
1	59.3	1490	69.65
550	63.4	2860	78.5

P	%	f.t.	P	%	f.t.
			E		
9	1	40.6	26	2000	55.0
21	1000	47.8	29	3000	61.5

P	t	S.C.*	P	t	S.C.*
	0 %			25 %	
1	39.7	12.3	1	42.8	13.5
1000	39.7	12.4	1000	48.7	13.7
2000	39.7	12.4			
	63.1 %			100 %	
1	49.6	20.4	1	59.3	33
1000	57.4	21.2	1000	66.3	34

S.C.* = speed of crystallization .

Zinc nitrate h. (ZnN_2O_6).6 H_2O
+ Cadmium nitrate h. (CdN_2O_6).4 H_2O

Pushin, 1949

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	60	-	45.5	36	28
85.5	56	-	31	28.5	28
77.5	53	-	23	31	27.5
62.5	46	27	0	36.4	-
53.5	41	27.5			

Calcium aluminum oxide h. ($Ca_3Al_2O_6$). 6 H_2O
+ Strontium Aluminum oxide h. ($Sr_3Al_2O_6$).6 H_2O

Carlson, 1955 (fig.)

Solid solutions

mol %	lattice constant Å	mol %	lattice constant Å
0	12.57	75	12.90
25	12.68	100	13.02
50	12.79		

Magnesium sulfate h. ($MgSO_4$).7 H_2O
+ Zinc sulfate h. ($ZnSO_4$).7 H_2O

Porter, 1930

%	t	n_D α (crystal axes)	n_D β (crystal axes)	n_D γ (crystal axes)
100	16-18	1.4567	1.4800	1.4840
90.76	19-19.5	.4538	.4773	.4812
77.44	17-19	.4502	.4739	.4781
64.95	18-20	.4474	.4710	.4750
54.92	18-19	.4450	.4682	.4720
42.93	16.5-19	.4425	.4658	.4699
33.30	18-20	.4402	.4630	.4674
27.79	18-20	.4385	.4620	.4659
18.69	18-19	.4364	.4598	.4645
6.10	20-20°5	.4339	.4568	.4619
0	18-20	.4321	.4551	.4605

Magnesium sulfate h. ($MgSO_4$).7 H_2O
+ Ferrous sulfate h. ($FeSO_4$).7 H_2O

Retgers, 1889

%	d	%	d
94.28	1.884	68.84	1.827
86.01	1.867	66.55	1.821
83.84	1.860	58.98	1.807
78.92	1.847	56.79	1.799
78.06	1.842	46.93	1.781

Magnesium sulfate h. ($MgSO_4$). 7 H_2O
+ Nickel sulfate h. ($NiSO_4$). 7 H_2O

Dufet, 1878

%	n_D	%	n_D
0	1.4554	71.95	1.479
28.35	.4645	79.1	.483
40.7	.4675	100	.4893
53.9	.472		

Magnesium ammonium sulfate h. ($MgN_2H_8S_2O_8$).6 H_2O
+ Ferrous ammonium sulfate h. ($FeN_2H_8S_2O_8$).6 H_2O

Retgers, 1890

%	d	%	d
100	1.865	38.23	1.770
90.20	.843	28.01	.760
82.24	.836	21.74	.747
67.78	.812	10	.731
65.44	.813	0	.721
48.58	.785		

Cadmium sulfate h. ($CdSO_4$). 7 H_2O
+ Ferrous sulfate h. ($FeSO_4$). 7 H_2O

Sommerfeldt, 1901

% (mixed crystals)	angle of optical axes
100.00	93.58
94.73	94.32
92.65	94.57
86.55	95.565

% (mixed crystals)	d
94.28	1.9186
90.66	.9425
86.56	.9619

Bismuth sulfate h. ($Bi_2S_3O_{12}$). 8 H_2O
+ Didymium sulfate h. ($Di_2S_3O_{12}$). 8 H_2O

Bodman, 1901

%	d	%	d
0	2.538	92.39	2.925-2.887
61.50	3.102-3.077	97.76	2.894-2.857
91.98	2.935-2.899	100.00	2.882

Bismuth sulfate h.($Bi_2S_3O_{12}$). 8 H_2O
+ Yttrium sulfate ($Yt_2S_3O_{12}$). 8 H_2O

Bodman, 1901

%	d	%	d
0	3.538	95.5	2.573
85.5	2.648	97.2	2.560
85.9	2.645	96.6	2.564
88.3	2.627	97.6	2.558
88.7	2.623	98.9	2.548
90.8	2.608	98.4	2.551
91.68	2.601	97.6	2.558
94.5	2.580	100.0	2.540
95.4	2.574		

Potassium alum ($KAlS_2O_8$).12 H_2O
+ Thallium alum ($TlAlS_2O_8$).12 H_2O

Dufet, 1874

mol %	n_D	mol %	n_D
0	1.45645	71.4	1.4847
9.3	.4602	78.4	.4867
14.3	.4627	85.8	.4926
32.7	.4700	88.0	.4927
57.8	.4764	100	.49748

Fock, 1880

mol%	n_D	mol%	n_D
0	1.4557	71.4	1.4847
9.3	.4602	78.4	.4847
14.3	.4627	85.8	.4926
32.7	.4700	88,0	.4927
57.8	.4764	100	.4888

Retgers, 1889

%	d	%	d
90.38	2.246	46.14	1.985
80.68	2.190	43.84	1.966
68.85	2.110	31.57	1.898
69.02	2.109	27.58	1.8777
66.30	2.070	24.82	1.864
52.87	2.015	15.46	1.821

Potassium aluminum alum ($KAlS_2O_8$).12 H_2O
+ Potassium chromium alum ($KCrS_2O_8$).12 H_2O

Rehberg, 1949

mol%	n	mol%	n
100	1.482	29.4	1.465
89.6	.480	22.1	.463
71.0	.475	11.3	.460
61.7	.473	4.6	.458
52.1	.470	0	.457
41.9	.468		

Strontium dithionate h. (SrS_2O_6).4 H_2O
+ Lead dithionate h. (PbS_2O_6).4 H_2O

Fock, 1880

mol%	n_D ordin.	n_D extraordin.
0	1.5296	1.5252
6.5	.5372	.5334
14.1	.5448	.5434
16.6	.5473	.5469
17.9	.5479	.5477
21.0	.5517	.5521
45.0	.5770	.5826
73.1	.6064	.6182
78.2	.6127	.6231
86.4	.6202	.6352
100	.6351	.6531

Lithium fluoride (LiF) + Tungsten anhydride (WO_3)

Schmitz-Dumont and al., 1953 (fig.)

mol%	f.t.	E	mol%	f.t.	E
35	800	760	10	800	-
25	760	760	0	830	-
20	770	740			

Sodium fluoride (NaF) + Tungsten anhydride (WO_3)

Schmitz-Dumont and al., 1953 (fig.)

mol%	f.t.	E	mol%	f.t.	E
50	690	-	20	840	540
40	780	550	0	980	-
30	810	550			

Potassium fluoride (KF) + Tungsten anhydride (WO_3)

Schmitz-Dumont and al., 1953 (fig.)

mol%	f.t.		E	
45	625		516	
42	516		"	
40	645		"	
25	950	(3+1)	745	tr.t.
20	940		724	
12	724		"	
10	790		"	
0	850		-	

Rubidium fluoride (RbF) + Tungsten anhydride (WO_3)

Schmitz-Dumont and al., 1953 (fig.)

mol%	f.t.	E	mol%	f.t.	E
40	590	610	12	685	685
25	918(3+1)	-	10	725	"
20	880	690	0	800	-

Cesium fluoride (CsF) + Tungsten anhydride (WO_3)

Schmitz-Dumont and al., 1953 (fig.)

mol%	f.t.	E
46	570	540
40	540	"
30	820	"
25	840 (3+1)	-
20	810	580 tr.t.
10	629	629
0	700	-

Calcium fluoride (CaF_2) + Aluminum oxide (Al_2O_3)

Pascal, 1913

%	f.t.	m.t.	%	f.t.	m.t.
0	1361	1361	30	1344	1304
20	1300	1270	100	2020	2020
25	1280	1270			

Sodium metaphosphate ($NaPO_3$) + Molybdenum anhydride (MoO_3)

de Carli, 1920

mol %	f.t.	mol %	f.t.
0	600	55	580
5	540	58	600
10	495	60	590
15	395	65	580
20	370	70	565
25	315	75	540
30	360	80	595
35	415	85	640
40	475	90	700
45	505	100	790
50	535		

(1+1)

Sodium metaphosphate ($NaPO_3$) + Tungsten anhydride (WO_3)

De Carli, 1926

%	f.t.	%	f.t.
0	600	40	460
5	540	45	505
10	500	50	530
15	450	55	600
20	400	60	790
25	385	65	850
30	350	70	840
35	410	75	765

(1+1)

Aluminum chloride ($AlCl_3$) + Antimony tribromide ($SbBr_3$)

Plotnikov and Schwarzman, 1938

mol%	f.t.	mol%	f.t.
100	96	50	161 tr.t.
92	86 E	40	173
90	93	30	180
80	118	20	182
70	148	10	188
60	153	0	190

Sodium (Na) + Sodium hydroxide (NaOH)

von Hevesy, 1909.

%	f.t.	%	f.t.
79.51	480	91.39	670
70.83	600	92.68	760
90.99	610	93.55	800

Potassium carbonate (K_2CO_3) + Niobium pentoxide (Nb_2O_5)

Reisman and Holtzberg, 1955

mol%	f.t.	m.t.	tr.t.	
0	891	-	-	
2	873	-	-	
3	870	-	-	
4	856	-	-	
5	851	793	-	
6	841	790	-	
7	834	794	-	
8	825	794	-	
9	816	795	-	
10	813	805	721	
11	799	794	-	
12	815	798	-	
13	830	795	-	
14	837	799	723	
15	840	790	726	
16	867	789	731	
17	889	781	727	
18	927	781	730	
19	936	779	"	
20	-	787	"	
21	-	783	731	
23	-	"	735	
24	-	-	"	
25	-	-	"	
28	947	845	-	
29	943	-	736	
30	939	839	-	
31	924	840	-	
32	910	830	-	
33	863	844	-	
35	879	"	-	
39	944	848	-	
40	957	844	-	
42.5	985	848	-	
45	1005	843	-	
50	1039	-	-	
51	1085	1039	-	
53	1120	"	-	
55	1141	1037	-	
56	1153	1042	-	(1+1)
57	1157	1039	-	
58	1159	1039	-	
59	1163	1025	-	
60	1163	-	-	(3+1)
62	1157	1147	-	
63	1156	1150	-	
65	-	1151	-	
67	1191	1151	-	
69	1208	1149	-	(2+3)
70	-	1150	-	
71	1224	1145	-	
73	1231	1141	-	
74	1232	1147	-	
75	-	-	-	(1+3)
76	1244	1213	-	
77	1253	1225	-	
78	1262	1232	-	
79	1270	1232	-	
80	1272	-	-	
82	1291	-	1279	
84	1331	-	1281	
85	1343	1234	1278	
86	1358	-	-	
87	1377	-	-	
89	1400	-	-	
90	1411	-	1280	
100	1486	-	-	

Titanium carbide (TiC) + Cobalt (Co)

Eremenko, 1956

%	f.t.	E
100	1490	-
98.2	1450	1360
97.1	1440	1350
95.0	1410	1370
94.0	1355	1355
92.5	1445	1370
89.7	-	1370
84.2	-	1365
82.8	-	1365

Titanium carbide (TiC) + Iron (Fe)

Eremenko, 1956

%	f.t.	E
100	1530	-
98.85	1518	-
97.50	1502	1450
96.37	1465-1455	-
95.75	-	1470
93.87	-	1460
93.10	-	1468
90.0	-	1462
87.3	-	1462
84.5	-	1452

XLIII. TWO SALTS WITH THE SAME METAL .

Lithium fluoride(LiF) + Lithium chloride (LiCl)

Botschwar, 1933 (fig.)

%	f.t.	E	%	f.t.	E
0	850	-	60	650	485
20	845	485	80	485	"
40	750	"	100	608	-

Lithium fluoride (LiF) + Lithium bromide (LiBr)

Botschwar, 1933 (fig.)

%	f.t.	E	%	f.t.	E
0	850	-	80	575	453
20	810	453	88	453	"
40	750	"	100	550	-
60	600	"			

Lithium fluoride(LiF) + Lithium hydroxide (LiOH)

Scarpa, 1915

wt%	mol%	f.t.	m.t.	E	min.
0	0	840	-	-	-
2.32	2.5	835	740	-	-
4.64	3	820	-	420	10
9.30	10	810	-	"	20
18.80	20	770	-	"	25
23.50	25	750	-	425	30
28.40	30	735	-	"	45
33.20	35	710	-	430	60
38.10	40	695	-	"	70
43.20	45	665	-	"	80
47.90	50	635	-	"	90
53.00	55	605	-	"	100
58.00	60	555	-	"	105
63.88	65	520	-	"	115
68.30	70	495	-	"	125
73.40	75	460	-	"	140
78.60	80	430	-	"	160
83.90	85	435	-	"	90
89.20	90	445	-	"	10
94.60	95	455	445	-	-
100	100	462	-	-	-

Lithium fluoride(LiF) + Lithium borate ($LiBO_2$)

Kitaigorodski, Popov and Botvinkin, 1933

wt%	f.t.	E	tr.t.			tr.t. complex	
			I	II	III	I	II
100	(845)	-	-	-	785	580	540
95	(825)	715	-	-	"	"	550
90	805	"	-	-		585	"
85	755	710	-	-	-	"	545
82	745	"	-	-	-	"	552
80	735	"	-	-	-	580	550
75	755	-	-	-	-	590	545
70	740	-	690	-	-	585	540
68	730	-	685	-	-	590	550
66	730	-	690	-	-	585	"
65	735	-	692	-	-	"	545
60	752	-	687	-	-	590	"
50	785	-	685	-	-	"	550
40	810	-	690	-	-	"	"
30	820	-	685	765	-	"	540
20	830	-	812	"	-	"	-
10	840	-	810	760	-	-	-
0	845	-	813	"	-	-	-
				(2+3)			

Lithium fluoride (LiF) + Lithium carbonate (Li_2CO_3)

Schmitz-Dumont and Heckmann, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	730	51	608 E
80	690	40	675
60	635	20	760
		0	850

Lithium fluoride (LiF) + Lithium silicate (Li_2SiO_3)

Bergman, Nesterova and Bichkova, 1955

mol%	f.t.	mol%	f.t.
0	843	21.4	960
0.8	846	23.5	972
1.6	844	25.9	982
2.4	842	27.4	994
3.1	850	29.8	1004
3.9	862	31.6	1014
5.3	874	34.3	1024
6.7	886	36.1	1032
8.2	894	38.9	1040
9.8	906	40.9	1050
11.8	916	43.5	1060
13.7	924	46	1070
15.7	936	49.6	1080
17.7	945	51.6	1088
19.8	950	53.9	1098

E : 840°

Lithium fluoride(LiF) + Lithium sulfate (Li_2SO_4)

Schmitz-Dumont and Heckmann, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	855	57	535 E
80	640	40	645
74	577 tr.t.	20	755
60	540	0	850

Lithium fluoride (LiF) + Lithium molybdate (Li_2MoO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	mol%	f.t.
100	702	40	700
80	655	20	780
62	617 E	0	830

Lithium fluoride (LiF) + Lithium tungstate (Li_2WO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	mol%	f.t.
100	740	50	675
80	690	40	710
62	642 E	20	770
		0	855

Kislova, Posipaiko and Bergman, 1955

%	f.t.	%	f.t.
100	738	38.0	690
90.5	724	33.7	712
81.9	710	29.1	730
74.0	695	25.9	750
66.7	680	21.2	766
60.0	664	17.7	778
53.9	650	14.3	792
51.6	644	11.2	802
48.2	638	8.2	812
46.0	648	5.3	822
40.9	674	2.6	834
		0	847

Lithium chloride (LiCl) + Lithium bromide (LiBr)

Botschwar, 1933 (fig.)

%	f.t.	%	f.t.
0	608	60	530
20	575	80	525
40	545	100	550

Lithium chloride (LiCl) + Lithium hydroxide (LiOH)

Scarpa, 1915

mol%	wt%	f.t.	tr.t.	min.	E	min.
0	0	605		-	-	-
5	2.87	575	305	20	-	-
10	9.40	540	310	40	-	-
20	12.10	500	315	60	-	-
30	19.5	400	"	100	-	-
40	28.2	380	"	130	-	-
45	31.6	340	"	60	280	20
50	36.1	315	"	-	"	40
55	40.8	308	-	-	290	100
60	45.9	300	-	-	"	140
65	51.2	300	-	-	"	120
70	56.8	330	-	-	"	90
80	69.3	375	-	-	285	70
90	83.5	425	-	-	280	50
95	91.4	445	-	-	"	30
100	100	460	-	-	-	-

Lithium chloride (LiCl) + Lithium metaborate ($LiBO_2$)

Posypaiko, Kislova and Bergman, 1956.

mol%	f.t.	mol%	f.t.
0	606	35	632
2.5	595	40	652
5	590 tr.t	45	664
7.5	587	50	682
10	585	55	700
15	580	60	718
17.5	576	70	750
19	572°E	80	774
20	576	90	805
22.5	585	92.5	815
25	596	94	815 tr.t
30	618	95	820
		97.5	830
		100	843

Lithium chloride (LiCl) + Lithium nitrate ($LiNO_3$)

Blidin, 1941

mol%	f.t.	mol%	f.t.
93.5	254	48.7	446
88.4	252	44.9	460
84.0	280	42.5	470
79.8	300	40.7	474
65.6	374	39.4	482
61.4	396	37.1	490
56.7	416	31.5	510
53.1	432		

E : 88.2 mol% 252°

Tokareva and Bergman, 1956.

mol%	f.t.	mol%	f.t.
100	255	82.2	277
97.7	252	79.5	297
96.0	250	77.2	310
94.6	248	75.0	325
91.2	246	71.2	345
88.3	245	67.0	362
87.5	244	57.5	406
84.5	258	54.5	418

Lithium chloride (LiCl) + Lithium vanadate (Li_3VO_4)

Bukhalov and Aleshkina, 1953 (fig.)

%	f.t.
0	606
1	598 E
7.7	800
18.2	900

Lithium chloride (LiCl) + Lithium sulfate (Li_2SO_4)

Dombrovskaya, 1933 (fig.)

mol%	f.t.	mol%	f.t.
100	845	38	482 E
67	572 tr.t.	25	520
42	500	11	505
		0	605

Dergunov, 1951

mol%	f.t.	mol%	f.t.
0	606	38	484
2.6	600	40.6	493
5.3	590	42.9	502
8.2	580	48.2	520
11.2	570	53.9	538
14.3	558	60	556
17.7	546	66.7	570
21.2	532	70.2	574
25.9	518	74	580
29.1	504	75.5	575 tr.t.
33.4	488	81.9	623
35.6	480	90.5	738
36	478 E	100	856

Evseiev and Bergman, 1951

mol%	f.t.	solid phase
0	605	I - LiCl
2.6	596	"
4.2	592	"
6.4	584	"
8.2	572 tr.t.	"
9.8	566	II - LiCl
11.8	563	"
14.3	560	"
17	550	"
19.1	542	"
21.2	538	"
26.6	523	"
32.5	496	"
36.1	486	"
37	480 E	
39	490	I - Li_2SO_4
50.4	522	"
54.7	540	"
63.3	556	"
72.5	570	"
74	572 tr.t.	I -
80.2	630	"
88.7	710	"

Lesnikh and Bergman, 1953

mol%	f.t.	mol%	f.t.
0	605	42.9	508
5.3	582	48.2	525
11.2	562	53.9	540
17.7	544	57.5	550
21.2	532	58.8	552
22	528	60	556
25	518	66.7	567
29.1	503	74	600
33.4	491	87	700
36.1	478	90.5	740
38	489	100	852

Bergman, Kislova and Posipaiko, 1954

mol%	f.t.	mol%	f.t.
0	606	37.0	482
3.1	593	38.0	488
6.4	582	40.9	494
9.9	566	43.9	503
14.3	555	49.3	519
18.4	538	54.7	538
23.5	526	62.7	552
25.9	520	69.5	568
27	515	75.5	590
29.1	509	78.6	610
30.8	501	81.9	628
32.5	495	87	695
34.3	490	90.5	740
35.6	482	100	852

transition : 10.5% 565°

: 77.8% 572°

Golubeva and Bergman, 1954

mol %	f.t.	mol %	f.t.
0	609	66.6	572 tr.t.
11.2	566 tr.t.	100	852
35.2	480 E		

Golubeva and Bergman, 1955

E : 34.5 mol % 485°

Lithium chloride (LiCl) + Lithium molybdate (Li_2MoO_4)

Lesnykh, Bergman and Bukun, 1956 (fig.)

mol%	f.t.	mol%	f.t.
0	600	33.4	510
11.2	558 tr.t.	38.0	518 tr.t.
14.3	550	60.0	625
25.9	498 E	40.0	700

Lithium chloride (LiCl) + Lithium tungstate (Li_2WO_4)

Lesnykh and Bergman, 1956.

mol%	f.t.
0	604
23.1	490 E
100	736

Bergman, Kislova and Posipaiko, 1954

mol %	f.t.	mol %	f.t.	mol %	f.t.
0	606	19.1	518	48.2	602
2.6	592	21.2	512	53.9	626
5.3	580	22.7	502	60	650
8.2	568	25.9	496	62.7	660
11.2	552	29.1	510	66.7	670
14.3	540	33.4	530	74	692
15.7	534	38	550	100	739
17.7	524	42.9	580		

E : 26.6 mol % 490°

tr.t.I : 9.2 mol % 565° tr.t.II : 63.3 mol % 660°

Lithium bromide (LiBr) + Lithium hydroxide (LiOH)

Scarpa, 1915

%	f.t.	tr.t.	min.	E	min.
0	550	-	-	-	-
1.43	520	-	-	250	10
3.00	485	-	-	255	30
6.50	440	-	-	275	50
8.41	400	-	-	"	60
10.7	380	-	-	"	80
16.5	320	-	-	"	120
18.4	275	-	-	"	150
22.0	285	-	-	270	120
25.2	295	-	-	275	100
29.3	300	-	-	265	80
33.9	310	-	-	270	50
39.2	330	310	60	265	20
45.3	355	"	125	-	-
52.5	380	"	100	-	-
71.8	425	305	50	-	-
83.9	440	"	20	-	-
100	462	-	-	-	-

Lithium iodide (LiI) + Lithium hydroxide (LiOH)

Scarpa, 1915

%	f.t.	tr.t.	min.	E	min.
0	440	-	-	-	-
0.93	405	-	-	180	30
1.90	368	-	-	"	40
3.7	340	-	-	185	60
4.4	312	-	-	180	70
5.6	290	-	-	"	90
7.1	245	-	-	"	110
8.7	230	-	-	"	130
10.6	205	-	-	"	150
12.8	180	-	-	"	180
15.2	230	-	-	182	140
17.9	265	-	-	"	120
21.1	295	-	-	"	100
24.9	310	-	-	180	60
29.4	315	-	-	165	40
34.9	328	-	-	160	20
45.7	340	330	120	-	-
50.3	375	"	90	-	-
61.7	410	"	60	-	-
77.2	440	325	30	-	-
100	462	-	-	-	-

Lithium metaborate ($LiBO_2$) + Lithium silicate (Li_2SiO_3)

van Klooster, 1910

%	f.t.	E	min.
100	1188	-	-
90	1150	-	-
80	1103	737 (804)	80
70	994	738 (803)	100
60	946	779 (802)	125
50	885	785 (802)	205
40	-	788 (802)	200
30	-	762 (803)	280
20	764	-	-
10	801	-	-
0	843	-	-

Lithium metaborate ($LiBO_2$) + Lithium tungstate (Li_2WO_4)

Kislova, Posipaiko and Bergman, 1955

%	f.t.	%	f.t.
100	738	33.7	776
95.1	718	31.2	780
90.5	710	29.1	782
86.1	694	27.0	786
81.9	694	25.0	790
77.8	704	23.1	794
74.0	710	21.2	798
70.2	710	19.5	800
66.7	716	17.7	804
63.3	724	16.0	808
60.0	730	14.3	810
56.9	740	12.7	812
53.9	744	11.2	814
51.0	750	9.6	814
48.2	754	8.2	814
45.5	760	6.7	814
42.9	762	5.3	820
40.4	765	3.9	828
38.0	770	2.6	836
35.6	773	1.3	840
		0	843

Lithium nitrate ($LiNO_3$) + Lithium carbonate (Li_2CO_3)

Amadori, 1913

%	f.t.	E	min.
0	255	-	-
2	-	250	800
5	420	250	800
10	502	248	750
15	542	245	680
20	582	245	540
30	614	236	420
50	-	240	300
75	-	230	180
100	732	-	-

Lithium nitrate ($LiNO_3$) + Lithium sulfate (Li_2SO_4)

Amadori, 1913

%	f.t.	E	min.	tr.t.
0	255	-	-	-
2	-	252	900	-
5	316	"	850	-
10	394	"	810	-
15	432	"	760	-
20	472	"	740	-
30	556	"	700	-
40	-	"	610	-
75	-	250	300	-
100	860	-	-	578

Lithium nitrate ($LiNO_3$) + Lithium acetate ($C_2H_3O_2Li$)

Diogenov, 1956.

mol%	f.t.	mol%	f.t.
0	259	56	170
5	254	60	183
9	248	62.5	190
14	241	68	206
18.5	234	72.5	219
22.5	227	77.5	233
27.5	216	81	242
32	204	84.5	250
37.5	188	90	259
41	176	92.5	264;263tr.t.
45	160	94	276
50	150	96	283
51	145°E	100	291
53	160		

Lithium metaphosphate ($LiPO_3$) + Lithium sulfate (Li_2SO_4)

Bergman and Sholokhovich, 1953 (fig.)

mol%	f.t.	mol%	f.t.
100	860	56.9	610
90.5	806	53.9	598
81.9	744	51.0	580
74.0	692	48.2	570
66.7	664	25	550
63.3	644	17.7	600
60.0	628	0	634

E : 35.6 tr.t. 50.0

Lithium carbonate (Li_2CO_3) + Lithium sulfate (Li_2SO_4)

Le Chatelier, 1897 (fig.)

mol%	f.t.	mol%	f.t.
100	830	50	580
90	745	40	605
80	667	30	630
70	580	20	650
65	550 E	10	675
60	560	0	695

Amadori, 1912

%	f.t.	tr.t.	min.
0	732	-	-
10	714	530	30
15	706	528	30
20	692	530	50
35	657	530	70
40	647	532	100
50	606	536	160
60	575	542	250
65		540	250
69.5	-	540	250
75	595	536	270
80	642	533	250
85	692	535	260
90	759	538	240
95	799	550	-
97.5	836	563	-
100	860	578	-

Belyaev and Doroshenko, 1956

mol%	f.t.	mol%	f.t.
100	700°	40	574
95	690	37.5	566
90	682	35.5	560 E
85	674	35	568
80	664	33	572 tr.t.
75	655	30	623
70	646	27.5	646
65	636	25	669
60	625	20	712
55	614	15	753
50	602	10	792
45	588	5	826
42.5	582	0	860

Lithium sulfate (Li_2SO_4) + Lithium molybdate (Li_2MoO_4)

Bergman, Kislova and Korobka, 1954

mol%	f.t.	mol%	f.t.
100	705	37.5	564 E
90	682	37	566
80	664	35	572
75	654	33.5	574
65	632	32.5	588
60	622	25	642
50	599	20	687
40	572	0	858

Lithium sulfate (Li_2SO_4) + Lithium tungstate (Li_2WO_4)

Bergman, Kislova and Posipaiko, 1954

mol %	f.t.	mol %	f.t.
100	738	64.4	634
98.7	734	62.0	624
97.4	730	59.6	620
96.1	724	57.1	614
94.7	720	54.5	608
93.3	716	51.8	600
91.8	712	48.5	596 E
90.4	705	46.2	614
88.8	702	43.1	634
87.4	696	40.0	652
85.7	690	36.7	674
84.0	685	33.3	690
82.3	680	29.8	710
80.5	674	26.0	734
78.8	666	22.2	756
76.9	660	18.9	776
75.0	656	13.9	800
73.0	652	9.5	818
70.9	648	4.9	842
68.8	644	0	860
66.6	636		

trans. : 23.1 mol % 660°

Lithium chromate (Li_2CrO_4) + Lithium tungstate (Li_2WO_4)

Kislova, Posipaiko and Bergman, 1955

%	f.t.	%	f.t.
100	738	45.0	638
95.0	726	40.0	626
90.0	714	35.0	610
85.0	704	30.0	594
80.0	696	25.0	576
75.0	690	20.0	556
70.0	684	15.0	540
65.0	678	10.0	530
60.0	670	5.0	520
55.0	662	0	517
50.0	650		

Sodium fluoride (NaF) + Sodium chloride (NaCl)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	820	50	770
90	785	40	810
80	755	30	850
70	720	20	900
66	695	10	950
60	635	0	980

Wolters, 1903

mol%	f.t.	E	min.
100	797	-	-
90	760	671	75
80	724	675	195
70	689	675	465
65.5	675	675	630
60	706	675	555
50	766	675	435
40	821	674	360
30	867	674	300
20	906	673	232
10	947	672	135
0	986	-	-

Plato, 1907

wt%	mol%	f.t.	E	min.
100.0	100	804.1	-	-
90	86.7	758.0	679.8	18.5
80	74.2	710.3	679.8	35.0
70	62.7	691.6	679.8	43.0
60	51.8	760.3	679.8	38.0
50	41.8	816.0	679.8	32.0
40	32.4	959.4	679.8	26.0
30	23.5	901.3	679.8	18.0
15	11.2	951.5	679.8	8.0
0	0	992.2	-	-

Volkov and Bergman , 1940 (fig.)

mol%	f.t.
100	800
65	676
60	820
0	990

Rasonskaya and Bergman , 1943 (fig.)

mol%	f.t.	mol%	f.t.
100	801	40	820
80	740	20	910
66.5	675	0	996
60	720		

Doucet, Bizouard and Netzer, 1956

m	t	m	t
0	799.85	4.620	720.61
0.534	789.85	5.150	713.69
0.765	786.00	5.387	709.08
1.395	772.85	5.910	700.61
1.894	766.00	6.360	697.54
2.381	758.31	6.281	690.61
2.881	749.08	7.375	684.46
3.500	738.31	7.780	678.31
4.050	729.08		

m = molality of NaF

Sodium fluoride (NaF) + Sodium bromide (NaBr)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	760	70	675
90	730	60	740
80	690	54	770
73	660	0	980

Sodium fluoride (NaF) + Sodium iodide (NaI)

Ruff and Plato, 1903

mol%	f.t.
100	680
90	645
80	615
74	650
0	980

Sodium fluoride (NaF) + Sodium hydroxide (NaOH)

Scarpa, 1915

wt%	f.t.	m.t.	E	min.	tr.t.I	II
0	1005	-	-	-	-	-
10	960	850	-	-	-	-
20	922	-	260	20	360	-
30	885	-	"	40	"	-
40	845	-	263	70	"	-
50	800	-	"	90	"	-
60	685	340	265	120	365	-
70	580	330	"	140	"	-
80	430	320	-	-	"	270
90	365	318	-	-	"	280
95	340	315	-	-	-	285
100	310	-	-	-	-	290

Sodium fluoride (NaF) + Sodium pyrophosphate ($Na_4P_2O_7$)

Sholokhovich and Belyaev, 1954 (fig.)

mol%	f.t.	mol%	f.t.
0	982	35	734 E
6	945	50	845
14	870	100	999
27	780		

Sodium fluoride (NaF) + Sodium carbonate (Na_2CO_3)

Amadori, 1913

%	f.t.	E	min.
100	854	-	-
95	808	684	50
90	764	688	90
85	724	690	160
80	-	690	200
71.59	744	690	180
65	784	690	170
60	815	690	120
50	856	690	110
40	902	685	80
20	956	678	50
10	982	674	20
0	1000	-	-

Volkov and Bergman , 1940 (fig.)

mol%	f.t	mol%	f.t
100	870	25	880
67	690	0	990
60	700		

Schmitz-Dumont and Heckman , 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	850	40	805
80	790	20	910
60	695 E	0	1000

Sodium fluoride (NaF) + Sodium metasilicate (Na_2SiO_3)

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	990	57 E	896
25	865	75	860
50	724	100	1088

Booth and Starrs, 1931

mol%	f.t.	E	mol%	f.t.	E
100	1089	-	42.39	926	912
86.84	1074	915	39.17	931	913
75.72	1049	913	33.48	947	"
66.29	1002	915	29.76	950	"
61.23	985	914	25.73	960	"
58.09	962	914	22.09	970	914
53.70	946	913	16.73	979	915
49.92	935	912	7.93	990	964
48.37	926	914	0	-	995
44.71	916	911			

Sodium fluoride(NaF) + Sodium titanate (Na_2TiO_3)

Sholokhovich, 1955

mol%	f.t.	mol%	f.t.
100	1025	42.9	907
94.2	1005	38.0	940
88.7	973	33.4	958
85.2	958	29.1	967
78.6	948	25.0	978
74	930	21.2	982
66.7	895	17.7	986
62.7	898	14.3	990
60	899	11.2	996
54.7	894	8.2	994
51.6	886	6.4	998
48.2	880	0	999

E_1 : 48.2 880° E_2 : 66.7 896°

Sodium fluoride (NaF) + Sodium disilicate ($Na_2Si_2O_5$)

Booth, Starrs and Bahnsen, 1933

mol%	wt%	f.t.	E
100	100	874	-
90	97.50	861	799
80	94.55	850	796
67.48	90.00	825	797
60.50	86.91	797	797
57.90	85.64	817	799
52.96	83.00	870	798
47.0	79.36	915	799
40.89	75.00	946	796
30.0	65.02	973	797
18.86	50.20	985	802
9.03	30.09	988	793
3.94	15.10	991	798
0	0	995	-

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	792	-	40	798	648
80	698	648	20	898	648
65	648	648	0	975	648
60	680	648			

tr.t. = 424°

Sodium fluoride (NaF) + Sodium sulfate (Na_2SO_4)

Wolters, 1903

mol%	f.t.	E	min.	tr.t.	min.
100	881	743	-	233	345
90	827	743	45	230	280
80	781	742	140	229	210
70	743	743	270	228	135
65	760	743	165	227	105
60	772	742	90	226	60
50	781	-	-	-	-
45	779	773	-	-	-
40	776	773	270	-	-
30	833	772	200	-	-
20	890	773	140	-	-
10	941	772	75 (1+1)	-	-
0	986	-	-	-	-

Gladushchenko and Bergman, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	990	50.4	784
14.3	915	70.2	747 E
35.2	735 E	100	884

(1+1)

Sodium fluoride (NaF) + Sodium chromate (Na_2CrO_4)

Rasonskaya and Bergman, 1943 (fig.)

mol%	f.t.	mol%	f.t.
100	780	40	880
81.5	642	20	940
60	790	0	996

Sodium fluoride (NaF) + Sodium molybdate (Na_2MoO_4)

Schmitz-Dumont and Weeg, 1951

mol%	f.t.	E
100	689	-
85	627	627
80	635	"
65	673	"
60	690	673
50	750	"
40	810	"
20	905	"
0	995	-

tr.t. - 620 582 426

Mateiko and Bukhalova, 1955

mol%	f.t.	mol%	f.t.
0	990	80	662
48	858	84	644
50	852	88	624
68	760	92	628
72	722	96	660
76	690	100	688

E : 89 mol% 611°

Sodium fluoride (NaF) + Sodium tungstate (Na_2WO_4)

Schmitz-Dumont and Weeg, 1951

mol%	f.t.	E	mol%	f.t.	E
100	692	-	50	740	620
80	635	635	40	805	618
70	670	"	20	905	-
61	693	"	0	990	-

tr.t. : 590°
574°
693°

Sodium chloride (NaCl) + Sodium bromide (NaBr)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	820	60	770
10	810	70	765
20	800	80	767
30	790	90	760
40	780	100	760
50	775		

Schobert, 1912

mol%	f.t.	mol%	f.t.
0	800	60	737
10	784	70	735
20	777	80	739
30	765	90	745
40	750	100	748
50	744		

Amadori, 1912

%	f.t.	%	f.t.
100	748	50	759
92.5	745	40	768
85	744	30	778
77.5	744	15	792
70	745	0	808
60	747		

Gromakov and Gromakova, 1955

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	800	800	60	745	739
10	788	775	70	741	739
20	777	759	80	739	738
30	767	749	90	739	738
40	758	742	100	742	742
50	751	740			

Wollam and Wallace, 1956

mol%	d (mixed crystals) pycnometric	d (mixed crystals) x-rays
	25°	
0.000	2.1615	2.1630
10.000	2.2829	2.2848
19.977	2.3971	2.3987
29.963	2.5069	2.5123
39.912	2.6169	2.6220
49.931	2.7203	2.7244
59.913	2.8255	2.8284
79.872	3.0160	3.0228
99.562	3.1980	3.1992

Nickels, Fineman and Wallace, 1949

mol%	molar volume (cc/mole)
25°	
10.18	27.543
20.19	28.081
29.82	28.559
40.29	29.105
49.27	29.569
59.77	30.113
70.07	30.627
79.22	31.115
89.47	31.605

Matsen and Beach, 1941

X - ray study of equimolar solid solutions

Fineman and Wallace, 1948

mol%	Q mix	mol%	Q mix
10.18	-136.7	59.77	-319.9
20.19	-220.5	70.07	-271.0
29.82	-291.9	79.22	-213.7
40.29	-323.5	89.47	-122.6
49.27	-335.5		

C.S.T. : 45 mol% 341° (mixed crystals)

Sodium chloride (NaCl) + Sodium iodide (NaI)

Greiner and Jellinek, 1933

mol% L	mol% V	p_1	p_2
	1180°		
0	0	73	0
25	52.2	57.0	62.2
50	77.7	39.8	138.9
75	90.2	22.2	194.0
100	100	-	260

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	820	60	605
10	795	65	600
20	760	70	605
30	730	80	630
40	690	90	655
50	630	100	680

Amadori, 1912

%	f.t.	E	min.
100	662	-	-
97	644-618	-	-
93	625-595	-	-
90	612	-	-
89.5	603	-	-
87	597	578	20
85	589	"	80
83	582	"	140
80		"	240
78	600	"	200
74	621	579	180
70	641	578	140
60	680	577	100
50	715	576	70
40	733	578	40
30	754	578	40
20	774	577	30
10	791	577	10
7	797	-	-
4	803	-	-
0	808	-	-

Schobert, 1926

mol%	f.t.	E	min.
0	800	-	-
2	795	-	-
5	786	566	10
10	770	569	60
20	740	572	170
30	705	572	240
40	667	571	280
50	627	574	390
60	579	571	490
70	586	569	220
73	598	570	120
80	613	587	-
90	645	(633)	-
100	670	-	-

Ilyasov and Bergman, 1956; Ilyasov and Bostandzhiyan 1956.

%	f.t.	%	f.t.
100	661	67.5	599
90	646	65	592
85	638	62.5	585 E
80	626	60	592
75	613	55	611
70	602	50	631

Sodium chloride (NaCl) + Sodium cyanide (NaCN)

Truthe, 1912

%	f.t.	m.t.	%	f.t.	m.t.
100	562.3	-	40	702	671
93	582	564	30	723	694
85	596	574	15.0	763	733
75	620	596	7	782	762
60	656	625	0	795	-
50	676	643			

Sodium chloride (NaCl) + Sodium hydroxide (NaOH)
Scarpa, 1915

%	f.t.	m.t.	tr.t. I	tr.t. II	E	min.
100	310	-	310	290	-	-
90	330	312	330	270	155	20
80	345	315	345	260	160	50
70	360	318	360	185	"	80
60	418	320	"	-	150	140
50	505	335	"	-	160	100
40	580	-	358	-	150	70
30	675	-	350	-	160	40
20	720	-	"	-	150	30
10	770	-	348	-	145	10
5	790	722	-	-	-	-
0	806	-	-	-	-	-

Antropoff and Sommer, 1926

%	f.t.	m.t.	tr.t. I	tr.t. II	E
100.0	322	-	-	303	-
95.0	325	321	-	289	182
90.0	326	321	-	277	178
85.0	330	325	-	252	190
80.0	335	329	-	237	169
70.0	356	340	-	186	173
60.0	360	342	-	-	165
50.0	504	347	360	275	170
40.0	580	348	360	319	159
30.0	644	-	358	-	159
20.0	699	-	357	-	167
10.0	752	-	357	-	164
0	798	-	-	-	-

Sodium chloride (NaCl) + Sodium nitrite ($NaNO_2$)

Meneghini, 1912

%	f.t.	%	f.t.
100	281.5	99	295
99.9	279-278.5	97-05	309
99.75	284-283.5	95	319
99.5	288.5	91.8	334.5
99.3	291		

Sodium chloride (NaCl) + Sodium phosphate (Na_3PO_4)

Le Chatelier, 1897

mol%	f.t.
0	778
10	762
20	743
30	710

Sodium chloride (NaCl) + Sodium carbonate (Na_2CO_3)

Le Chatelier, 1897 (fig.)

mol%	f.t.	mol%	f.t.
0	778	60	635
10	755	67	615
20	738	70	635
30	718	80	690
40	692	90	755
50	665	100	810

Amadori, 1913

%	f.t.	E	min.
100	854	-	-
95	828	626	50
87.5	780	632	80
75	710	634	160
66	-	636	210
64.46	670	636	190
50	662	636	170
40	702	636	150
30	728	632	90
10	785	632	40
0	808	-	-

Niggli, 1919

mol%	f.t.	E	mol%	f.t.	E
0	801	-	53.5	664	639
38	720	638	59	645	641
41	700	638	79	774	639
42	699	638	100	860	-

Volkov and Bergman , 1940 (fig.)

mol%	f.t.
0	800
26	750
45	638
100	860

Ryschkewitsch, 1933

t	κ (mhos)	t	κ (mhos)
0 mol%		20 mol%	
805	3.54	700	2.27
820	.62	740	.51
833	.68	770	.87
830	3.35	800	.90
840	.40	810	.93
880	.525	850	.95
930	.735	870	3.32
940	.965	900	.32
960	4.02	850	.35
		1000	.41
		1020	.63
		1040	.73
		1060	.74

t	κ	t	κ
33 mol%		100 mol%	
640	1.30	830	1.60
650	.58	835	.64
660	.85	850	2.37
680	2.21	860	.58
690	.17	875	.64
720	.20	900	.69
750	.51	960	.76
800	.665	1000	.84
850	3.07	1020	.92
875	.07	1030	.82
900	.04	1050	.84
915	.04		
930	.07		
935	2.98		
950	3.10		
960	3.13		
970	3.13		
1000	3.10		
1010	3.20		
1050	3.20		

Sodium chloride (NaCl) + Sodium metatitanate (Na_2TiO_3)

Sholokhovich and Barkova, 1956 (fig.)

mol%	f.t.	mol%	f.t.
100	1023	76.3 - 1.6	933 $C+L_1+L_2$
98.1	796 E	0	800

Sodium chloride (NaCl) + Sodium sulfate (Na_2SO_4)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	820	45	650
10	795	50	660
20	760	60	705
30	725	100	890
40	680		

Wolters, 1903

mol%	f.t.	E	min.	tr.t.
100	881	-	-	233
90	833	622	67	224
80	787	623	120	223
70	738	622	195	222
60	691	623	270	221
50	639	623	330	220
45	625	624	345	219
40	650	624	290	218
30	698	624	225	217
20	725	623	165	216
10	761	623	90	215
0	797	-	-	-

Jänecke, 1908

mol%	f.t.	E	tr.t.
100	897	-	225
60	700	630	225
48.2	-	628	230
33.4	690	630	225
14.3	770	630	230
0	810	-	-

Flood, Förland and Nesland, 1951

mol%	f.t.
100.0	885.0
93.91	857.5
88.70	835.3
81.85	802.7
58.01	683.2

Akopov and Bergman, 1954

mol%	f.t.	mol%	f.t.
100	884	25.0	716
66.0	724	17.0	748
59.5	691	10 5	762
53.1	661	4.8	785
48.2	628	0	800
35.2	679		

E : 48.2 mol% 628°

Bergman and Bakumskaya, 1955

mol%	f.t.	mol%	f.t.
0	800	50.5	654
5.3	790	56.9	690
14.3	760	63.3	724
25.0	722	66.7	740
33.4	694	70.2	752
38.0	676	74	774
40.4	664	77.8	794
42.9	654	81.9	820
45.5	638	100	884
48.2	628		

Sodium chloride (NaCl) + Sodium chromate (Na_2CrO_4)

Rasonskaya and Bergman, 1943 (fig.)

mol%	f.t.	mol%	f.t.
0	801	42.9	640
11.2	770	65.3	572
25.0	715	100	780

Gromakov, 1951

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	780	780	50	592	566
90	740	740	40	644	564
80	700	566	30	685	564
70	654	564	20	730	-
60	606	564	10	765	-
54	564	-	0	800	800

Sodium chloride (NaCl) + Sodium molybdate (Na_2MoO_4)

Bukhalova and Mateiko, 1955

mol%	f.t.	mol%	f.t.
0	800	50.5	643
17.5	749	57.5	642
25	722	62.5	640
31.5	690	67	(628)
36	666	74	613
38	650	82	635
43	640	100	688
47	643		

Sodium chloride (NaCl) + Sodium tungstate (Na_2WO_4)

Bukhalova and Mateiko, 1956.

mol%	f.t.	mol%	f.t.
100	698	49.3	680(1+1)
90.5	666	46	680
81.9	644	39.9	672
81.3	637 E	37	666
77	644	35.6	662 E
72.5	652	34.3	670
68	660	31.6	682
60	670	26.6	704
56.3	675	0	800

Sodium bromide (NaBr) + Sodium iodide (NaI)

Amadori, 1912

%-	f.t.	%	f.t.
100	662	60	655
92.5	652	45	674
85	648	30	696
77.5	645	15	721
70	647	0	748

Schobert, 1912

mol%	f.t.	mol%	f.t.
0	748	50	648
10	722	60	640
20	699	70	637
30	679	80	644
40	663	100	670

Sodium bromide (NaBr) + Sodium hydroxide (NaOH)

Scarpa, 1915

wt%	mol%	f.t.	E	min.	tr.t.
0	0	765	-	-	-
2.5	6.19	750	260	20	-
10	22.32	675	"	40	-
20	40.00	575	"	70	-
30	52.44	475	"	90	-
40	63.29	395	262	110	-
50	71.84	320	260	135	-
60	79.36	260	"	150	-
70	85.78	275	"	120	-
80	90.91	290	255	70	-
90	95.74	302	250	20	290
100	100	310	-	-	"

Sodium bromide (NaBr) + Sodium nitrite ($NaNO_2$)

Meneghini, 1912

●%	f.t.	%	f.t.
100	281.5	99	283.5
99.9	278-277.5	97.5	293
99.75	278	96	302
99.5	280-280.5	95	307
99.25	282		

Sodium bromide (NaBr) + Sodium carbonate (Na_2CO_3)

Gromakov, 1951

mol%	f.t.	E	mol%	f.t.	E
0	740	-	50	665	608
10	708	608	60	730	608
20	673	608	70	778	-
30	650	608	(80)	806	-
40	619	608	(90)	830	-
(43)	608	-	100	852	-

Sodium bromide (NaBr) + Sodium sulfate (Na_2SO_4)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	760	40	645
10	735	50	675
20	705	100	890
30	670		

Flood, Förland and Nesland, 1951

mol%	f.t.
90.52	843.1
82.31	804.9

Sodium iodide (NaI) + Sodium hydroxide (NaOH)

Scarpa, 1915

%	mol%	f.t.	tr.t.	min.	E	min.
0	0	665	-	-	-	-
5	16.34	615	290	20	-	-
10	29.41	560	"	40	220	20
12	33.77	540	295	50	"	30
15	40.10	505	"	70	"	40
20	48.54	430	300	40	"	70
30	61.98	320	"	20	-	-
35	67.05	290	-	-	225	100
40	71.43	285	-	-	"	120
45	75.75	260	-	-	"	140
50	79.10	235	-	-	"	150
60	84.74	245	-	-	"	120
70	89.74	270	-	-	220	90
80	93.89	290	-	-	215	40
90	96.98	305	285	-	210	20
100	100	310	290	-	-	-

Sodium iodide (NaI) + Sodium nitrate ($NaNO_3$)

Vasenin and Bergman, 1938

%	f.t.
86 E	296
100	308

Sodium iodide (NaI) + Sodium sulfate (Na_2SO_4)

Nyankovskaya, 1956

mol%	f.t.	mol%	f.t.
0	660	40	642
3	657	45	660
6	651	50	681
9	647	55	700
12	644	65	739
15	637	70	763
20	628	75	777
25	608	80	797
29.5	598 E	85	817
30	600	90	837
35	618	100	884

Sodium thiocyanate (NaCNS) + Sodium nitrate ($NaNO_3$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	308	50	236
90	288	40	253
80	268	30	269
70	252	20	285
60	236	10	299
55	232 E	0	311

Sodium thiocyanate (NaCNS) + Sodium formate ($NaCHO_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	258	55	212
95	250	50	232
90	241	45	244
85	233	40	256
75	213	35	267
70	202	25	284
65	190	10	302
64	187 E	0	311
60	197		

Sodium thiocyanate (NaCNS) + Sodium acetate ($NaC_2H_3O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	331	50	256
95	326	45.5	244 E
90	320	45	245
85	313	40	258
75	302	35	266
70	295	25	282
65	287	10	302
60	278	0	311
55	268		

Sodium thiocyanate (NaCNS) + Sodium propionate ($NaC_3H_5O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	298	50	265
95	292	46	258 E
90	289	45	261
85	287	40	269
75	282	35	276
70	280	25	285
65	276	10	298
60	273	0	311
55	270		

Sodium thiocyanate (NaCNS) + Sodium butyrate ($NaC_4H_7O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	330	55	263
95	328	51.5	262 E
90	324	50	269
85	316	45	280
75	291	40	287
70	275	35	290
68.5	268	25	298
65	266	10	304
60	264	0	311

Sodium thiocyanate (NaCNS) + Sodium isobutyrate ($NaC_4H_7O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	260	55	266
95	247	50	274
90	237	45	280
85	231	40	284
75	221	35	288
72.6	214 E	25	295
70	221	10	300
65	240	0	311
60	255		

Sodium thiocyanate (NaCNS) + Sodium valerate ($NaC_5H_9O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	357	54	289
95	370	50	291
90	378	45	288
85	375	43.5	287 E
75	356	40	290
70	344	35	293
65	331	25	297
60	366	10	302
55	296	0	311

Sodium thiocyanate (NaCNS) + Sodium isovalerate ($NaC_5H_9O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	262	55	285
95	277	50	291
90	287	45	297
85	288	40	300
75	270	35	303
70	256	25	305
68	250 E	10	306
65	260	0	311
60	272		

Sodium thiocyanate (NaCNS) + Sodium caproate ($NaC_6H_{11}O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	365	50	330
95	383	45	316
90	397	40	305
85	399	37	295 E
75	386	35	297
70	377	25	304
65	366	10	307
60	352	0	311
55	342		

Sodium sulfide (Na_2S) + Sodium carbonate (Na_2CO_3)

Tammann and Oelsen, 1930

%	f.t.	E	%	f.t.	E
100	855	-	59	750	755
90	830	755	50	845	"
80	800	"	40	930	"
70	775	"	30	980	"
60	750	"	0	1040	-

Sodium sulfide (Na_2S) + Sodium sulfate (Na_2SO_4)

Grube, 1927

%	f.t.	%	f.t.
100	885	60	765
90	845	50	811
80	808	40	857
70	769		

E : 67% 748°

Tammann and Oelsen, 1930

%	f.t.	E	min.
100	888	-	-
90	842	740	67
84.5	812	"	102
78.2	762	"	152
72.5	-	"	194
65.2	780	"	185
60	802	"	161
50	851	"	131
40	910	"	110
0	1040	-	-

Sodium hydroxide ($NaOH$) + Sodium nitrate ($NaNO_3$)

Retortillo and Moles, 1933

mol%	f.t.	mol%	f.t.
0	320	49.6	239
6.25	298	49.6	271.5
9.4	321	49.82	260
12.5	298	51.0	260.5
19.7	287.5	54.0	260-251
22.5	262	62.5	252.5
23.0	264	59.0	255.5
25.3	250	70.19	240
29.4	254	70.19	240.5
32.5	262-260	79.8	250
34.1	260	84.5	260
39.7	253.5	85.6	263
45.2	257.5	100	307.5
49.26	260		

(2+1) (1+1)

Sodium hydroxide ($NaOH$) + Sodium arseniate (Na_3AsO_4)

Urazov, Lovchikov and Lipshits, 1956, (fig.)

%	f.t.	%	f.t.
0	330	40	480
10	300	42	485
20	400	50	600
30	435	55	650

Sodium hydroxide ($NaOH$) + Sodium carbonate (Na_2CO_3)

Maksimenko, 1913

%	f.t.
0	321
13	300 E
100	850

Neumann and Bergve, 1914

%	f.t.	%	f.t.
0	290	2.55	290.0
2.7	290.0	6.48	284.0
12.4	284.7	12.36	280.5
17.2	280.0	16.30	280.0
24.5	283.0	24.40	284.0
29.9	297.0	28.57	298.5

Seward, 1942

%	f.t.	tr.t.	E
0	320	-	-
0.4	319	294	-
2.4	316	294	286
4.3	313	294	-
7.1	307	294	286
9.7	301	294	-
12.2	295	-	286
14.6	292	-	286
19.0	288	-	286
22.8	292	-	286
25.3	320	-	-
26.6	333	-	286
27.9	342	-	-
29.4	354	-	-
30.8	367	-	-
33.1	381	-	286
35.1	395	-	-

Sodium nitrite ($NaNO_2$) + Sodium nitrate ($NaNO_3$)

Bruni and Meneghini, 1909

%	f.t.	%	f.t.
0	284.5	48.3	222
1.5	281.5	50.75	221.5
3.95	278	53.2	223
3.4	274	55.65	225
11.45	267	58.1	230
16.3	259	60.6	235
21.2	251	70.45	249
31.05	233	80.3	267
36	226	90.15	288
40.9	224.5	95.07	297
43.35	224	100	312
45.8	223		

%		%	
L	C	L	C
21.2	14.7	43.35	38.3
33.5	26.0	55.65	56.0
36.0	29.2	70.45	74.4
38.4	35.1	85.2	88.15

Voskresenskaya, Jankovskaya and Anosov, 1948

	U (cal/g)			U (cal/g)	
t	C	L	t	C	L
	45.1 %			0 %	
110.0	0.303	-	110.0	0.320	-
150.7	.333	-	152.7	.366	-
179.6	.451	-	200.5	.413	-
213.8	.464	-	242.2	.431	-
234.8	.581	0.429	260.0	.446	-
271.0	.555	.409	290.5	.616	0.482
310.2	.529	.387	320.0	.593	.408
350.5	.509	.380	356.5	.569	.383
492.5	.458	.359			

Sodium carbonate (Na_2CO_3) + Sodium tetraborate ($Na_2B_8O_{13}$)

Burgers and Holt, 1905

mol %	f.t.	mol %	f.t.
83.4	917	50.0	664
71.5	722	45.5	692
62.5	783	41.7	685
55.6	654		

Leontyeva, 1936

t	d	t	d	t	d
		80 mol %			
1000	2.012	850	2.123	690	2.252
950	2.049	800	2.160	600	2.326
900	2.088	750	2.203		
		50 mol %			
1000	2.079	900	2.141	800	2.203
950	2.109	850	2.169	790	2.208
900	2.137				
		33.4 mol %			
1010	2.000	900	2.075	750	2.212
1000	2.004	850	2.114	700	2.232
950	2.041	800	2.151		
		21 mol %			
1000	1.992	845	2.088	800	2.132
950	2.020	830	2.109	700	2.174

Sodium borate ($NaBO_2$) + Sodium metaphosphate ($NaPO_3$)

van Klooster, 1910

%	f.t.	%	f.t.
100	610	30	864
80	608	20	866
70	796	10	922
60	800	0	966
50	774		

Sodium borate ($NaBO_2$) + Sodium phosphate (Na_3PO_4)

Le Chatelier, 1894 and 1897

%	f.t.	%	f.t.
0	940	41	960 (1+1)
7	918	44	950
9	910	50	930
23	932	71.5	850
33	952	89.5	925
37.5	960	100	970

Sodium borate ($NaBO_2$) + Sodium metasilicate (Na_2SiO_3)

van Klooster, 1910

%	f.t.	E	min.
100	1056	-	-
90	1011	759	20
80	940	806	115
70	-	815	245
60	-	815	260
50	-	814	300
40	870	782	160
30	891	760	80
20	924	737	60
10	938	754	10
0	966	-	-

Sodium tetraborate ($Na_2B_4O_7$) + Sodium sulfate (Na_2SO_4)

Zhilenko and Sverchkov, 1940

mol %	f.t.	sat.t.	mol %	f.t.	sat.t.
5	glass	-	50	862	L_1+L_2
10	"	-	60	864	"
15	"	-	70	866	"
17	700	-	80	868	"
18	723	-	85	870	"
20	760	-	90	872	970
25	850	896	92.5	880	900
30	864	970	95	882	-
35	864	1020	100	884	-
40	864	L_1+L_2			

Sodium aluminate ($NaAlO_2$) + Sodium ferrite ($NaFeO_2$)

Toropov and Shishakov, 1939

	biaxial refractive index 5890 Å	
%	Ng	Np
0	1.580 ± 0.002	1.566 ± 0.002
50.02	.781 ± 0.002	1.751 ± 0.002
80.54	.965 ± 0.01	.905 ± 0.01
100	2.08 ± 0.01	2.01 ± 0.01

Sodium chlorate ($NaClO_3$) + Sodium bromate ($NaBrO_3$)

Huber, 1940 (fig.)

mol%	lattice constant	mol %	lattice constant
0	6.555	75	6.650
25	.585	100	.690
50	.610		

Sodium nitrate ($NaNO_3$) + Sodium carbonate (Na_2CO_3)

Amadori, 1913

%	f.t.	E	min.
0	310	-	-
3	-	304	500
5	328	304	480
7.5	404	304	440
10	446	304	420
15	494	304	380
20	548	302	340
25	602	300	320
35	658	290	260
50	-	286	190
75	-	278	110
100	854	-	-

tr.t. = 430°

Sodium nitrate ($NaNO_3$) + Sodium sulfate (Na_2SO_4)

Amadori, 1913

%	f.t.	E	min.
0	310	-	-
5		300	300
10	320	"	320
15	376	"	300
20	420	"	280
25	480	"	270
30	534	298	240
35	558	298	220
40	608	296	200
50	-	298	170
60	-	298	140
80	-	280	60
100	892	-	-

tr.t. = 238°

Sodium nitrate ($NaNO_3$) + Sodium sulfamate ($NaSO_3NH_2$)

Laning and Van der Meulen, 1948

%	f.t.	E	%	f.t.	E
100	250.0	-	58.33	205.3	-
85	226.5	-	57.5	205.2	-
80	219.0	-	56.67	205.2	205.0
75	208.3	198.3	55	208.6	205.0
72.5	202.5	198.8	50	220.0	-
70	201.6	199.0	45	231.6	-
66.7	203.8	-	44.6	232.0	-
65	204.7	-	40.2	238.8	-
60	205.3	-	0	307.5	-
58.35	205.7	-			

E_1 : 71.5% 199°

E_2 : 56.6 205°

(1+1) 205.7°

Sodium nitrate ($NaNO_3$) + Sodium chromate (Na_2CrO_4)

Rasonskaya and Bergman, 1953

%	f.t.	%	f.t.
0	308	6.4	332
1.6	308	8.2	360
2.6	308	11.2	390
3.1	307	14.3	412
3.7	304	17.7	456
4.2	302	21.2	485
4.8	300	29.1	542
5.3	316		

Sodium nitrate ($NaNO_3$) + Sodium formate ($NaCHO_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	258	51	186 E
95	252	50	190
90	242	45	206
85	232	40	220
75	214	35	235
70	206	25	262
65	198	15	284
60	192	5	302
55	188	0	308

Sodium nitrate ($NaNO_3$) + Sodium acetate ($NaC_2H_3O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	331	50	247
95	328	45	233
90	324	42	224 E
85	318	40	231
75	304	35	242
70	296	25	264
65	286	15	284
60	276	5	304
55	263	0	308

Diogenov, 1956.

mol%	f.t.	mol%	f.t.
100	337	40.5	230
97	323	38.	235
93	321	34	245
91.5	319	30.3	255
85.5	312	29.5	257
81	306	24.5	266 (1+2)
75	297	21.6	270
73	292	20	272 (4+1)
70.5	287	18.5	270
66.7	278	16.5	270
63.5	270	14.5	278
55.5	263	10	290
52	257	8.5	294
50	253	5.5	299
45	240	3.3	304
42.5	225 E	0	308

Sodium nitrate ($NaNO_3$) + Sodium propionate ($NaC_3H_5O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	298	50	264
95	294	45	258
90	291	43.5	255 E
85	287	40	261
75	282	35	270
70	280	25	280
65	276	15	290
60	273	5	301
55	269	0	308

Sodium nitrate ($NaNO_3$) + Sodium butyrate ($NaC_4H_7O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	330	55	268
95	329	50	267 E
90	325	45	276
85	315	40	283
75	283	35	288
73	276	25	296
70	274	15	298
65	273	5	302
60	271	0	308

Sodium nitrate ($NaNO_3$) + Sodium isobutyrate ($NaC_4H_7O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	260	50	274
95	248	45	276
90	242	40	280
85	238	35	284
75	219 E	25	288
70	233	15	292
65	244	5	300
60	248	0	308
55	267		

Sodium nitrate ($NaNO_3$) + Sodium valerate ($NaC_5H_9O_2$)

Sokolov, 1954

mol%	f.t.		mol%	f.t.
100	357		55	294
95	366		50	295
90	372		45	288
85	369	$L_1 + L_2$	41.5	281 E
75	350		40	285
70	336		35	293
65	320		25	298
60	296		15	300
59.5	295		5	305
			0	308

Sodium nitrate ($NaNO_3$) + Sodium isovalerate ($NaC_5H_9O_2$)

Sokolov, 1954

mol%	f.t.		mol%	f.t.
100	262		55	282
95	273		50	288
90	280		45	294
85	280	$L_1 + L_2$	40	299
75	265		35	302
70	257		25	304
68.8	254 E		15	306
65	260		5	307
60	272		0	308

Sodium nitrate ($NaNO_3$) + Sodium caproate ($NaC_6H_{11}O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
100	365	55	326
95	376	50	313
90	383	45	296
85	385	43.5	287 E
75	375	41	300
70	367	40	302
65	357	0	308
60	342		

39-0.16 % $L_1 + L_2$

Sodium metaphosphate ($NaPO_3$) + Sodium pyrophosphate ($Na_4P_2O_7$)

Parravano and Calcagni, 1909

%	f.t.	E	min.
44	-	595	110
50	-	612	150
54	682	612	130
64	768	612	105
70	839	-	-
80	921	-	-
90	962	-	-
100	988	-	-

Partridge, Hicks and Smith, 1941

%	f.t.	%	f.t.
0	625	45	622
10	617	50	643
20	597	60	780
30	572	72	865
35	579	100	980
40	610		

Morey and Ingerson, 1944

%	f.t.	tr.t.	E	
0	627.6	-	-	A
15	607	-	552	"
33	557	-	552	(1+1)
40	586.5	-	552	"
45	603	-	-	"
47.5	615.5	-	-	"
50	631	622	-	β
56.59	704	622	552	"
57	711	622	-	"
60	743	-	-	"
66	801	-	-	"
72	880	622	-	"
(100)	(985)	-	-	-

Sodium metavanadate ($NaVO_3$) + Sodium metatitanate (Na_2TiO_3)

Sholokhovich and Barkova, 1956 (fig.)

mol%	f.t.	mol%	f.t.
100	1023	50	1046 (1+1)
85	950	25	900
79.5	920 E	0.5	638 E
65	980	0	640

Sodium pyrophosphate ($Na_4P_2O_7$) + Sodium metatitanate (Na_2TiO_3)

Sholokhovich and Barkova, 1956 (fig.)

mol%	f.t.	mol%	f.t.
100	1023	20	1040
90	982	10	936
80	1080	0	999
	(C+L_1+L_2)		

Sodium pyrophosphate ($Na_4P_2O_7$) + Sodium molybdate (Na_2MoO_4)

Belyaev and Sholokhovich, 1953

mol%	f.t.	mol%	f.t.
100	690	70.9	819
98.4	684	68.4	819
96.9	674	66.6	820
95.2	663	64.8	816
94.1	663	63.0	815
93.0	684	60.1	826
91.8	714	58.1	836
90.4	732	50.7	865
88.2	756	46.1	884
86.3	774	40.0	906
84.3	786	33.3	926
82.3	796	26.0	946
78.8	810	18.1	966
75.0	818	9.5	982

E_1 : 94.7 654° E_2 :63 807°

Sodium pyrophosphate ($Na_4P_2O_7$) + Sodium tungstate (Na_2WO_4)

Sholokhovich and Bergman, 1954

mol%	f.t.	mol%	f.t.
100	694	70.9	788
96.9	676	68.4	808
95.2	666	66.3	818
93.6	656	63.9	830
91.8	647 E	60.1	848
90.1	675	56.1	865
88.2	706	52.9	880
86.3	720	49.6	890
84.3	730	46.1	898
82.3	740	37.3	910
80.2	751	36.0	928
78.0	758	30.5	960
75.7	764	0	999
73.4	767		
		(1+1)	

Sodium carbonate (Na_2CO_3) + Sodium sulfate (Na_2SO_4)

Le Chatelier, 1894 and 1897

mol%	f.t.	mol%	f.t.
100	860	33	790
67	810	22	795
50	800	0	820

Amadori, 1912

%	f.t.	tr.t.	%	f.t.	tr.t.
0	854	430	88	866	-
10	848	-	95.5	874	-
20	837	-	95.5	882	183
35	828	-	96.3	886	200
47	828	-	98.5	888	223
62	833	-	100	892	238
75	848	-			

Sodium carbonate (Na_2CO_3) + Sodium chromate (Na_2CrO_4)

Vilnyanski and Pudovkina, 1947

%	f.t.	E	%	f.t.	E
0	855	-	62.5	655	655
10	830	-	70	670	-
20	800	650	80	700	650
30	775	650	90	740	-
40	745	660	100	792	-
50	700	660			

Sodium carbonate (Na_2CO_3) + Sodium aluminosilicate ($NaAlSiO_4$)

Eitel, 1925 (fig.)

%	f.t.	E	%	f.t.	E
0	864	-	60	1095	843
2	843	843	80	1170	-
20	915	"	90	1200	-
40	1010	"	100	1526	-

Burgess and Holt, 1903

mol%	f.t.	mol%	f.t.
16.6	917	54.5	692
37.5	783	58.3	685
44.4	654	28.5	722
50.0	664		

Sodium metasilicate (Na_2SiO_3) + Sodium sulfate (Na_2SO_4)

Tammann and Oelsen, 1930

%	f.t. II	f.t. I	E
100	886	-	-
97	877	-	868
94	871	-	867
90	863	881	868
80	849	902	867
70	834	921	867
60	819	931	868
50	799	945	861
48	770	965	857
30	-	983	848
20	-	1018	841
10	-	1040	835
0	-	1086	-

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
100	884	40	1010
97	870	20	1040
80	955	0	1088
60	985		

Sodium metasilicate (Na_2SiO_3) + Sodium chromate (Na_2CrO_4)

Vilnyanski and Pudovkina, 1947 (fig.)

%	f.t.	E	%	f.t.	E
100	792	-	60	985	765
90	770	770	50	-	760
80	855	770	40	-	745
70	930	765			

Sodium silicate (Na_2SiO_3) + Sodium tungstate (Na_2WO_4)

van Klooster, 1910

%	f.t.	E	tr.t. I	tr.t. II
100	700	-	589	572
99	762	699	590	578
97	850	688	591	570
94.5	937	687	590	576
91	1008	690	592	569
16.6	1050			
0	1056			

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
100	698	75	1005
99	695	50	1020
97	915	25	1050
90	975	0	1088

Sodium metasilicate (Na_2SiO_3) + Sodium disilicate ($Na_2Si_2O_5$)

Tilley, 1932

%	f.t.	%	f.t.
0	1089	73.37	862
30.33	1040	79.28	851
52.94	955	81.85	858
60.66	920	90.98	868
64.30	904	100	874

E : 76.7% 846°

Sodium metasilicate (Na_2SiO_3) + Sodium metagermanate (Na_2GeO_3)

Schwartz and Lewinsohn, 1930

%	f.t.	m.t.	%	f.t.	m.t.
100	1078	-	40	1075	1062
90	1074	1059	30	1075	1060
80	1073	1060	20	1075	1070
70	1070	1059	10	1083	1080
60	1062	1060	0	1087	-
50	1071	1062			

Sodium metatitanate (Na_2TiO_3) + Sodium chromate ($NaCrO_4$)

Sholokhovich and Barkova, 1956, (fig.)

mol%	f.t.
0	1023
6.5-98.0	967 $C+L_1+L_2$
99.5	802 E
100	966

Sodium metatitanate (Na_2TiO_3) + Sodium tungstate (Na_2WO_4)

Sholokhovich and Barkova, 1956, (fig.)

mol%	f.t.
0	1023
2.5 - 93.0	961 $C+L_1+L_2$
99.5	696 E

Sodium metatitanate (Na_2TiO_3) + Sodium sulfate (Na_2SO_4)

Sholokhovich and Barkova, 1956, (fig.)

mol%	f.t.
0	1023
5 - 97	967 $C+L_1+L_2$
98.5	866 E
100	884

Sodium metatitanate (Na_2TiO_3) + Sodium molybdate (Na_2MoO_4)

Sholokhovich and Barkova, 1956, (fig.)

mol%	f.t.	mol%	f.t.
0	1023	99.5	682 E
6.0-97.0	977 $C+L_1+L_2$	100	687

Sodium sulfate (Na_2SO_4) + Sodium molybdate (Na_2MoO_4)

Belyaev and Doroshenko, 1954 (fig.)

mol%	f.t.	mol%	f.t.
100	700	25	790
75	673	0	884
50	705		

Boecke, 1906

mol%	f.t.	E	tr.t.				
			I-II	I-V	II-III	III-IV	
100	692	-	619	-	-	587	431
98	690	689	602	-	-	-	230
90	684	684	-	548	540	-	-
80	680	680	-	510	503	-	-
70	688	680	-	574	468	-	-
60	682	682	-	524	410	-	-
50	608	680	-	373	-	-	-
40	723	689	-	322	-	-	-
30	758	689	-	-	-	-	-
22.8	-	-	-	231	-	-	-
20	802	686	-	212	212	-	-
10	844	-	-	229	212	-	-
0	888	-	-	239	-	-	-

Sodium sulfate (Na_2SO_4) + Sodium tungstate (Na_2WO_4)

Boecke, 1906

mol%	f.t.	m.t.	tr.t:				
			II-III	I-II	I-III	I-V	III-IV
100	698	-	588	-	-	-	564
95	685	683	601	-	-	-	525
90	672	669	608	-	-	-	486
89	669	665	608	-	-	-	472
88	663	663	-	625	597	558	-
85	663	663	-	-	560	536	-
80	662	662	-	-	533	520	-
75	662	662	-	-	511	498	-
70	662	662	-	-	488	458	-
65	663	663	-	-	453	-	-
60	666	666	-	-	-	-	-
50	679	664	-	-	-	-	-
40	708	666	-	-	-	-	-
30	731	668	-	-	219	-	-
20	789	669	-	-	230	222	-
10	850	666	-	-	239	-	-

Sodium acid sulfate ($NaHSO_4$) + Sodium pyrosulfate ($Na_2S_2O_7$)

Cambi and Bozza, 1923

wt%	f.t.	tr.t.	min.
100	400.9	-	-
90.13	392.7	173.0	4
81.30	380.0	"	2
77.68	374.2	162.0	3
49.97	326.5	176.5	17
16.82	245.0	180.0	15
11.01	221.5	182.3	16
6.43	185.0	182.7	20
4.37	183.0	"	21
3.91	183.2	"	16
3.64	185.0	182.0	10
2.26	185.1	-	-
-	185.7	-	-

Sodium molybdate (Na_2MO_4) + Sodium tungstate (Na_2WO_4)

Boecke, 1906

mol%	f.t.	m.t.	tr.t. I	II	III	IV	V	VI
0	692	-	619	-	587	-	431	-
10	685	685	629	625	579	579	435	-
20	684	684	645	640	580	580	455	-
30	684	684	670	-	581	581	470	-
35	683	683	-	-	581	581	478	-
40	684	684	-	-	581	581	484	475
50	688	683	-	-	581	581	503	494
60	691	686	-	-	582	582	516	505
70	691	686	-	-	576	576	530	522
80	693	689	-	-	572	572	544	543
90	695	693	-	-	575	575	555	550
100	698	-	-	-	-	-	654	-

Sodium formate (CHO_2Na) + Sodium acetate ($C_2H_3O_2Na$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	258	55	300
5	252	60	303
10	244	65	308
10.5	242 E	70	313
15	252	75	316
20	260	80	320
25	267	85	323
30	270	90	326
35	278	95	330
40	284	100	331
45	291		
50	296		

Sodium formate (CHO_2Na) + Sodium butyrate ($C_4H_7O_2Na$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	258	55	340
2.5	252 E	60	340
5	287	65	339
10	301	70	338
15	312	75	336
20	318	80	331
25	324	85	324
30	327	89	308 E
35	333	90	311
40	337	95	322
45	339	100	330
50	341 (1+1)		

Sodium formate (CHO_2Na) + Sodium isobutyrate ($C_4H_7O_2Na$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	258	55	329
1.3	252 E	60	327
5	290	65	325
10	305	70	320
15	314	75	314
20	319	80	306
25	321	85	296
30	324	90	282
35	326	95	258
40	327	96.5	250 E
45	329	100	260
50	330 (1+1)		

Sodium formate (CHO_2Na) + Sodium isovalerate ($C_5H_9O_2Na$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	258	50	317
0.75	252	55	312
5	287	60	309
10	300	65	306
15	308	70	301
20	311	75	297
25	314	80	284
30	316	90	265
35	318	94.5	245 E
40	320 (3+2)	100	262
45	319		

Sodium acetate ($NaC_2H_3O_2$) + Sodium butyrate ($NaC_4H_7O_2$)

De Mol, 1925 (fig.)

mol%	f.t.	E	mol%	f.t.	E
0	330	-	39	261.5	261.5
10	313.	261.5	47.5	280	261.5
20	292.5	261.5	50	-	261.5
30	274	261.5			

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	331	55	265
5	319	60	260
10	309	65	254
15	299	69	250 E
20	290	70	253
25	282	75	266
30	274	80	281
33.5	266 E	90	312
35	268	95	324
40	273 (3+2)	100	330
45	270		
50	268		

Sodium acetate ($NaC_2H_3O_2$) + Sodium isobutyrate ($NaC_4H_7O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	331	55	218
5	323	58	208 E
10	314	60	215
15	305	65	226
20	297	70	236
25	288	75	242
30	277	80	246
35	265	85	250
40	254	90	254
45	242	95	257
50	230	100	260

Sodium acetate ($NaC_2H_3O_2$) + Sodium isovalerate ($NaC_5H_9O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	331	55	215
5	320	60	199
10	311	65	184
15	304	70	166
20	295	73	156 E
25	287	75	163
30	280	80	185
35	269	85	207
40	260	90	228
45	248	95	247
50	232	100	262

Sodium acetate ($NaC_2H_3O_2$) + Sodium caproate ($NaC_6H_{11}O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	331	49.5	260 E
5	321	50	265
10	312	55	300
15	304	60	321
20	296	65	332
25	288	70	342
30	279	75	349
34.5	268 E	80	353
35	269	90	360
40	269	95	363
45	265	100	365

(5+3)

Sodium acetate ($NaC_2H_3O_2$) + Sodium benzoate ($NaC_7H_5O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	331	20	411
2.6	315 E	25	421
5	350	30	428
10	380	33	431
15	400	100	465

Sodium butyrate ($NaC_4H_7O_2$) + Sodium isobutyrate ($NaC_4H_7O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	330	55	235
5	317	60	229
10	306	65	224
15	297	70	222
20	287	72.5	221 E
25	279	75	222
30	270	80	225
35	264	85	228
40	257	90	235
45	250	95	248
50	242	100	260

Sodium butyrate ($NaC_4H_7O_2$) + Sodium isovalerate ($NaC_5H_9O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	330	55	287
5	326	60	284
10	323	65	281
15	320	70	277
20	316	75	273
25	312	80	269
30	308	85	263
35	305	90	258
40	300	90.5	257 E
45	295	95	263
50	292	100	262

Sodium butyrate ($NaC_4H_7O_2$) + Sodium caproate ($C_6H_{11}O_2Na$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	330	50	334
5	328	55	340
10	324	60	343
15	320	65	344
20	318	70	349
22.5	317 E	75	353
25	321 (3+1)	80	356
27.5	317 E	85	359
30	319	90	360
35	323	95	363
40	326	100	365
45	331		

Sodium butyrate ($NaC_4H_7O_2$) + Sodium benzoate ($NaC_7H_5O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	330	35	401
0.13	330 E	40	408
5	349	45	415
10	361	50	421
15	370	55	427
20	378	60	434
25	386	100	463
30	394		

Sodium isobutyrate ($C_4H_7O_2Na$) + Sodium isovalerate ($C_5H_9O_2Na$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	260	55	189
5	248	60	191
10	238	65	194
15	229	70	199
20	220	75	207
25	213	80	215
30	207	85	225
35	201	90	237
40	197	95	248
45	193	100	262
50	188		

Sodium isobutyrate ($NaC_4H_7O_2$) + Sodium caproate ($NaC_6H_{11}O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	260	45	191
5	237	50	205
10	218	55	220
15	195	60	235
20	175	65	252
23.5	160 E	70	270
25	161	75	290
30	168	85	329
35	175	90	345
40	182	95	356
		100	365

Sodium isobutyrate ($NaC_4H_7O_2$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	260	50	319
5	240	55	321
10	215	60	323 (2+3)
15	196	65	322
20	177	70	321
25	163	75	320
25.5	162 E	85	317
30	217	90	314
35	260	94.5	311
40	291	97.5	312 (sic)
45	309	100	308

Sodium isobutyrate ($NaC_4H_7O_2$) + Sodium benzoate ($NaC_7H_5O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	260	45	344
3.5	228 E	50	355
5	235	55	367
10	256	60	379
15	272	65	389
20	288	70	399
25	301	75	408
30	312	100	463
35	322		
40	335		

Sodium isovalerate ($NaC_5H_9O_2$) + Sodium caproate ($NaC_6H_{11}O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	262	55	283
5	255	60	291
10	247	65	296
15	243	70	304
20	239	75	313
25	242	80	322
30	248	85	331
35	256	90	341
40	263	95	354
45	271	100	365
50	277		

Sodium isovalerate ($NaC_5H_9O_2$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	262	55	305
5	201	60	315
10	167	65	322
15	147	70	323
17.3	140 E	75	321
20	162	85	318
30	225	90	314
35	247	93.5	309
40	266	95	312
45	282	100	308
50	295		

(1+2)

Sodium isovalerate ($NaC_5H_9O_2$) + Sodium benzoate ($NaC_7H_5O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	265	40	389
3	261 E	45	396
5	275	50	401
10	298	55	407
15	317	60	415
20	337	65	421
25	249	70	426
30	365	100	463
35	379		

Sodium caproate ($NaC_6H_{11}O_2$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	365	55	326
5	320	60	329 (2+3)
10	272	65	328
15	242	70	327
17.5	239 E	75	326
20	248	80	324
25	264	85	321
30	280	90	319
35	293	94.5	314
40	305	95	316
45	313	100	308
50	320		

Sodium caproate ($NaC_6H_{11}O_2$) + Sodium benzoate ($NaC_7H_5O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	365	20	380
5	369	35	396
10	373	40	400
13	371 E	100	463
15	374		

Sodium laurate ($NaC_{12}H_{23}O_2$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Vold, 1941

mol%	tr.t. subneat-neat	tr.t. superwaxy-subneat I	tr.t. superwaxy-subneat II	tr.t. waxy-superwaxy
0	244	220	-	182
13.2	249	213	184	169
20	249	214	184	163
25.3	247	-	-	171
38.2	247	214	186	165
42.5	247	224	192	165
50	246	221	193	169
50	246	218	197	171
62.5	249	214	192	164
71	253	-	196	169
74.8	253	-	201	167
84.6	257	-	204	171
94.9	253	-	207	170
100	253	208	208	172

mol%	tr.t. subwaxy	tr.t. waxy	tr.t. waxy-superwaxy	tr.t. superwaxy subneat	tr.t. subneat -neat
94.9	113	132-137	171	203	252
94.6	109	131-134	172	205	-
75	100	127-132	167	201-211	260
62.5	96	136-145	-	198-222	-
50	92	135	164	192-211	-
25.3	101	125-132	-	-	-

mol %	tr.t.*	mol %	tr.t.*
0	336	50	310
5.2	334	62.5	307
13.2	329	74.8	308
25.3	324	84.6	305
38.2	320	94.9	301
42.5	313	100	295
50	312		

*tr.t. isotropic-liquid neat soap .

Ferguson and Nordsieck, 1944 (fig.)

mol %	long spacing (Å)	mol %	long spacing (Å)
	type I		
0	32	23	32.3
5	32.2	28	32.5
15	32.3		
	type II		
37	36.5	87	39.4
48	36.6	90	39.5
55	37	92	39.6
70	38.5	100	40
77	38.8		

Sodium laurate ($NaC_{12}H_{23}O_2$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Vold, 1941

mol%	tr.t. neat-isotrope	tr.t. neat-subneat	tr.t. subneat-superwaxy I	tr.t. subneat-superwaxy II	tr.t. super-waxy-waxy	tr.t. waxy-curd
88.9	289	255	-	205	166	126
77.9	293	255	209	201	171	-
66.5	296	259	222	198	169	114
59.5	298	253	224	199	170	-
46.7	307	253	224	199	167	116-130
37.8	311	252	224	195	169	115-132
25.7	315	-	-	-	-	-
19.4	317	249	220	192	-	-
11.7	322	-	-	-	-	-
8.2	329	243	218	194	171	147

Ferguson and Nordsieck, 1944 (fig.)

mol%	long spacing (Å)
type I	
0	32
10	32
15	32
23	32
type II	
63	43.5
75	44
90	44
100	45
type III	
23	38.5
35	38.5
42	38.5
53	38.5
63	38.5

Sodium myristate ($NaC_{14}H_{27}O_2$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Ferguson and Nordsieck, 1944 (fig.)

mol%	long spacing (Å)
0	36
15	36.5
27	36.5
38	37.5
48	38
52	38.2
69	38.2
85	39.5
100	40

Sodium palmitate ($NaC_{16}H_{31}O_2$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Vold, 1941

phase change	100	74.8	50.1	25.7	0
			tr.t. mol%		
subwaxy	117	-	-	-	117
waxy	132	126	126	-	135
waxy-superwaxy	167	164	166	172	172
superw-subneat	203	202	206	207	208
subneat-neat	257	253	253	253	253
neat-isotrope	288	289	293	295	295

Sodium palmitate ($NaC_{16}H_{31}O_2$) + Sodium behenate ($NaC_{22}H_{43}O_2$)

Vold, 1941

phase change	13.8	33.0	57.4	82.7
		tr.t. mol%		
neat-isotrope	-	287	275	272
subneat-neat	254	256	258	257
superw-subneat	204	203	205	200
waxy	128	-	120	125
superwaxy-waxy	170	163	-	-

Sodium stearate ($NaC_{18}H_{35}O_2$) + Sodium oleate ($NaC_{18}H_{33}O_2$)

Vold, 1941

%	neat-isotrope	subneat-neat	superwaxy-subneat
	tr.t.		
91.5	271	257	207
73.6	278	252	193
49.0	285	254	194
23.6	288	256	201

Sodium stearate ($NaC_{18}H_{35}O_2$) + Sodium benzoate ($NaC_7H_5O_2$)

Sokolov, 1954

mol%	f.t.	mol%	f.t.
0	308	30	362
1.3	301 E	35	369
5	310	40	376
10	321	45	384
15	332	50	390
20	344	55	396
24	353	100	463

Sodium potassium tartrate ($NaKC_4H_4O_6 \cdot 4H_2O$) + Sodium ammonium tartrate ($NaNH_4C_4H_4O_6 \cdot 4H_2O$)

Eremeyev and Kurchakov, 1933

Mixed crystals

mol%	Curie point	mol%	Curie point
0	24	21.5	-70
1	5	53.5	-100
3	0	59	-106
11.5	-15	75	-167

Potassium fluoride (KF) + Potassium chloride (KCl)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	795	60	650
90	755	54	624
80	725	50	630
70	690	0	885

Plato, 1907

wt%	mol%	f.t.	E	min.
100	100	772.3	-	-
90	87.6	733.4	605.1	12.5
80	75.7	692.6	605.1	22.5
70	64.5	648.1	605.1	34.0
60	53.9	612.1	605.1	44.0
50	43.8	667.1	605.1	39.0
40	35.0	713.1	605.1	32.0
30	25.0	757.1	605.1	24.0
20	16.3	795.4	605.1	17.0
10	8.0	830.0	605.1	9.0
0	0	859.9	-	-

Potassium fluoride (KF) + Potassium bromide (KBr)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	750	60	615
90	720	50	650
80	685	45	680
70	645	0	885
61	610		

Kurnakov and Wrzesnewsky, 1912

mol%	f.t.	E	pressure of flow
		15.16°	(kg/mm²)
100	748	-	40.5
97	740	582	-
95	736	583	46
90	720	580	50.2
80	680	580	62
70	645	582	71.8
60	580	580	79
50	637	582	88
30	754	584	96.5
10	835	581	113.2
5	843	580	-
3	847	581	-
0	851	-	124.6

Potassium fluoride (KF) + Potassium iodide (KI)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	705	70	580
90	670	60	615
80	625	0	885

Potassium fluoride (KF) + Potassium hydroxide (KOH)

Scarpa, 1915

wt%	mol%	f.t.	m.t.	tr.t.
0	0	857	-	-
10	10.40	835	790	-
20	20.34	795	750	-
30	30.63	750	700	138
40	40.80	710	660	142
50	51.14	665	615	155
60	61.14	625	560	170
70	71.02	570	505	188
80	80.79	505	460	200
90	90.39	450	415	220
95	95.15	420	395	235
100	-	380	-	265

Potassium fluoride (KF) + Potassium pyrophosphate ($K_4P_2O_7$)

Sholokhovich and Belyaev, 1954 (fig.)

mol%	f.t.	mol%	f.t.
0	865	26.9	712 E
6.0	835	30.0	800
7.6	812	(12+1)50.0	970
14.3	790	100	1099

Potassium fluoride (KF) + Potassium carbonate (K_2CO_3)

Amadori, 1913

%	f.t.	E	min.
100	876	-	-
95	852	678	40
90	812	684	70
85	768	688	90
80	735	688	140
75	-	688	220
70.38	688	-	-
67.50	688	-	-
65	686	-	-
60	-	682	280
50	722	682	150
40	768	682	90
25	812	682	60
10	840	680	30
0	855	-	-

Potassium fluoride (KF) + Potassium metasilicate (K_2SiO_3)

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	852	63 E	730
25	828	80	835
50	685	100	960

Potassium fluoride (KF) + Potassium titanate (K_2TiO_3)

Sholokhovich, 1955

mol%	f.t.	mol%	f.t.
0	870	33.4	788
2.6	868	38.0	776
5.3	860	43.9	764
8.2	852	48.2	777
11.2	844	53.9	782
14.3	835	60	790
17.7	826	66.7	798
21.2	817	71.0	806
25.0	806	81.9	811
29.1	798	100	826

E : 33.4 mol% 758°

Potassium fluoride (KF) + Potassium dititanate ($K_2Ti_2O_5$)

Schmitz-Dumont and Schulz, 1952 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	940	-	12	846	810
75	925	825	6	849	750
50	895	820	0	850	-
20	820	820			

Potassium fluoride (KF) + Potassium sulfate (K_2SO_4)

Karandeeff, 1909

%	f.t.	E	min.	tr.t. I	tr.t. II	min.
100	1074	-	110	599	-	140
96.4	1034	880	28	598	-	125
92.3	987	882	72	599	-	112
87.5	938	881	120	596	588	42
81.8	887	882	246	596	588	10
81.3	883	883	278	595	586	10
75.0	887	-	234	-	578	-
66.7	878	786	92	-	578	-
56.3	855	786	176	-	587	-
42.9	804	789	377	-	584	-
37.9	788	788	480	-	585	-
24.8	819	786	225	-	-	-
0	867	-	330	-	-	-

Benrath and Drek, 1921

t	$\varkappa.10^6$ 0	10	20 mol%	30	40	50
555	-	-	0.316	0.646	0.417	0.195
580	-	0.170	0.603	0.977	0.708	0.516
605	-	0.363	1.07	1.78	1.18	0.724
630	0.178	0.631	1.66	2.51	1.78	1.07
655	0.324	1.00	2.40	3.72	2.69	1.59
679	0.617	1.55	3.63	5.75	3.98	2.29
704	1.15	2.46	5.75	8.51	6.03	3.47
728	2.04	4.07	8.71	12.6	9.33	5.01
752	3.80	6.31	13.5	18.6	14.8	7.59
776	7.24	372	275	315	309	11.0
799	12.6	-	-	-	-	16.6
823	24.6	-	-	-	-	24.6
846	46.8	-	-	-	-	37.2
869	4170	-	-	-	-	60.3
892	-	-	-	-	-	550

t	60	70	80 mol%	90	100
555	-	-	-	-	3.16
580	0.724	1.95	0.562	0.661	4.47
605	1.55	4.68	2.09	2.63	6.31
630	2.40	6.61	3.09	4.17	8.91
655	3.55	9.33	4.90	6.61	12.6
679	5.13	12.9	7.41	10.0	17.4
704	7.94	18.6	11.2	15.5	24.0
728	12.0	25.7	17.8	24.6	33.9
752	17.8	36.3	27.5	36.3	47.9
776	27.5	52.5	42.7	57.5	69.2
799	41.7	74.1	64.6	89.1	100
823	66.1	107.2	102	138	126
846	105	166	155	214	174
869	159	275	257	324	240
892	10000	22900	14100	5750	339

Potassium fluoride (KF) + Potassium chromate (K_2CrO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	976	-	40	740	732
80	885	758	30	732 E	732
60	785	758	20	760	732
53	758 E	758	0	855	-
50	762	-			

tr.t. 670° tr.t. (1+1) 616°

Potassium fluoride (KF) + Potassium molybdate (K_2MoO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	936	-	40	748	722
80	840	749	30	722 E	"
60	749 E	"	20	760	"
50	752	-	0	855	-
tr.t.	475				
	439				
	321				
(1+1)	350				

Mateiko and Bukhalova, 1955

mol%	f.t.	mol%	f.t.
0	858	46	724
4	848	50	734
10	833	54	743
15	822	57	748
18	812	62	750
24	790	66	754
28	780	70	750
32	770	74	770
36	756	78	804
40	742	100	926
43	728		

E_1 : 45 mol% 722°

E_2 : 72 mol% 745°

(2+1) : 754°

Potassium fluoride (KF) + Potassium tungstate (K_2WO_4)

Schmitz-Dumont and Weeg, 1951

mol%	f.t.	E	mol%	f.t.	E
100	922	-	40	750	729
80	835	757	38	729 E	"
60	757 E	"	20	790	"
50	761	"	0	855	-

tr.t. = 370° (1+1)

Kislova, Posipaiko and Bergman, 1955

mol%	f.t.	mol%	f.t.
0	852	29.1	734
2.6	838	31.2	740
5.3	824	33.4	746
8.2	812	35.2	750
9.3	804	38	758
11.2	796	39.9	760
12.4	790	42.9	763
14.3	783	48.2	764
15.7	776	53.9	762
17.7	766	60.0	780
19.2	760	66.7	805
21.2	752	74.0	838
22	745	81.9	868
25.0	736	90.5	896
26.0	728	100	926

(1+1)

Potassium chloride (KCl) + Potassium bromide (KBr)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	790	60	745
10	785	70	740
20	770	80	745
30	760	90	750
40	750	100	750
50	748		

Amadori and Pampanini, 1911

%	f.t.	%	f.t.
100	740	40	745
90	736	30	752
80	734	20	758
70	735	10	767
60	737	0	774
50	739		

Wrzesnewsky, 1912

mol%	f.t.	m.t.	pressure of flow after mixing	after 5 weeks
0	781	-	28.1	-
10	771	760	57.5	-
20	758	744	71	54
30	743	731	75.2	-
40	733	723	76.6	63.5
45	724	720	-	-
50	719	717	73.9	-
60	716	716	70.1	61.3
70	719	716	-	-
80	723	716	58.0	-
90	730	727	48.6	-
100	748	-	36.1	36.0

Wollam and Wallace, 1956

mol%	d (mixed crystals) pycnometric	x-rays
	25°	
0.000	1.9868	1.9862
20.00	2.1558	2.1602
30.00	2.2377	2.2386

Boardman, Palmer and Heymann, 1955 (fig.)

mol %	σ 750°	800°
0	102	98
20	98	95
40	95	92
60	93	89
80	91	87
100	90	86

Matsen and Beach, 1941

X-ray study of equimolar solid solutions at different temperatures

Zhemchuzhni and Rambach, 1910

mixed crystals

mol %	Q mix (cal/g)
75	785
50	820
25	565
after 1 1/2 month	
50	610

Popov, Bundel and Choler, 1930

mixed crystals

Q mix

20 mol %	40 mol %	60 mol %	80 mol %
mecanic mixture,powder			
- 46	-	- 63	- 47
rapid crystallization by heating			
-126	-210	-205	-103
-144	-210	-211	-142
slow crystallization			
-165	-210	-210	-154
slow crystallization with subsequent heating and slow cooling			
-172	-235	-225	-148

Fontell, 1938-40

mol%	Q mix cal/g	cal/mole
75	1.482	160
67.5	1.967	206
49.7	2.397	232
37.5	2.334	213
25	2.160	185
12.5	1.412	113

Mc Coy and Wallace, 1956.

mol%	Q mix (mixed crystals)
25°	
10	70
20	135
30	185
40	215
50	222
60	217
70	196
80	157
90	95

Eucken and Kuhn, 1928

mol%	thermic conductivity . 10^5 0°	-190°
90	694	1185
75	467	673
50	566,552	606,612
25	804	1058
10	1155	1859
0	2102	1433

Williams, 1956 (fig.)

mol %	Thermal resistivity mol % (cm-deg/watt) -268°	-258.2°
0	0.08	0.5
10	1.39	3.5
20	2.04	5.25
30	2.63	6.75
40.	3.04	7.7
50	3.25	8.3

Potassium chloride (KCl) + Potassium iodide (KI)

Greiner and Jellinek, 1933

mol% L	V	p_1	p_2
1180°			
0	0	106	0
25	40.3	80.8	54.5
50	67.1	57.1	117.2
75	88.1	23.6	172.3
100	100	0	231

Le Chatelier, 1898

mol%	f.t.	mol%	f.t.
100	640	33	630
83	610	20	680
67	590	0	740
50	580		

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	790	53	640
10	770	60	645
20	745	70	655
30	720	80	675
40	685	90	690
50	650	100	705

Ilyasov and Bergman, 1956.

mol%	f.t.	mol%	f.t.
0.0	772	55.0	584
10.0	746	60.0	587
20.0	705	70.0	599
30.0	653	80.0	638
40.0	600	90.0	668
50.0	586	100.0	680
52.5	584 E		

Amadori and Pampanini, 1911

%	f.t.	E	min.
100	680	-	-
95	659-639	-	-
90	636-621	-	-
85	623-610	-	-
80	609-603	-	-
75	602	-	-
72	600	-	-
70	600	600	300
65	622	600	250
60	640	600	160
50	672	598	80
40	698	599	50
30	721	598	30
20	741	597	10
18	745	-	-
15	750-668	-	-
10	758-726	-	-
0	774	-	-

Wrzesnewzky, 1912

mol%	f.t.	m.t.	tr.t.	pressure of flow
0	780	-	-	28
10	745	732	597	50.8
20	707	676	578	58.5
30	652	623	548	30
40	595	592	510	17.8
45	590	588	498	10.1
50	586	584	496	8.8
55	584	584	496	8.6
60	588	584	498	9.8
70	599	591	505	17.0
80	646	615	504	44.0
90	668	650	-	20.3
100	680	-	-	-

Tammann and Ruppelt, 1931

Mixed crystals

mol%	sat.t. by cooling beginn.	by cooling end	by heating beginn.	by heating end
10	530	465	490	590
20	570	515	550	600
30	600	555	550	605
50	530	500	520	560
70	455	430	445	515

Van Artsdalen and Yaffe, 1935

t	d	t	d
	0%		
779.5	1.5219	870.4	1.4688
781.8	.5211	907.2	.4479
809.4	.5058	939.0	.4292
827.4	.4943		
	19.78%		
711.8	1.7825	823.8	1.7056
728.2	.7719	847.0	.6890
728.5	.7716	870.2	.6735
751.2	.7550	904.6	.6501
775.2	.7392		
793.7	.7253		
	38.88%		
679.5	1.9856	829.0	1.8730
694.6	.9744	855.1	.8535
736.6	.9426	881.9	.8327
784.1	.9069	911.9	.8103
805.9	.8900		
	54.85%		
618.8	2.1742	806.7	2.0131
650.2	.1474	809.6	1.9991
665.2	.1348	874.4	.9621
690.2	.1135	902.3	.9397
735.7	.0755	902.7	.9394
735.8	.0754	902.7	.9394
762.3	.0537		
	74.33%		
640.3	2.3033	790.3	2.1698
675.4	.2718	818.7	.1443
700.6	.2499	846.6	.1198
728.4	.2244	879.0	.0906
729.1	.2237	914.1	.0595
756.9	.1994		
	84.04%		
706.0	2.3536	874.2	2.2081
763.0	.3055	901.2	.1856
820.9	.2540		
	93.96%		
682.1	2.3962	845.4	2.2440
726.6	.3550	869.5	.2215
778.6	.3063	904.4	.1900
843.5	.2739		
	100%		
681.9	2.4467	789.9	2.3434
706.3	.4242	809.4	.3245
730.1	.4009	819.9	.3145
745.0	.3865	825.6	.3090
752.3	.3795	847.6	.2887
767.7	.3645	847.7	.2886
786.5	.3472	878.6	.2590
		903.4	.2355

t	κ(mhos)	t	κ(mhos)
	0%		
778.9	2.1628	850.9	2.3609
782.4	.1816	851.2	.3618
787.0	.1984	865.0	.3901
790.0	.2074	877.5	.4174
799.0	.2342	889.0	.4476
800.5	.2407	892.5	.4484
804.1	.2478	905.5	.4745
823.0	.3020	906.7	.4777
846.0	.3507	925.1	.5156
	19.78%		
711.8	1.7043	814.9	1.9905
726.8	.7530	817.8	.9998
741.6	.7936	845.6	2.0559
745.4	.8071	865.0	.0843
761.2	.8517	875.1	.0985
767.4	.8712	883.3	.1130
790.2	.9243	903.1	.1402
797.6	.9500		
	38.88%		
642.3	1.0824	732.6	1.6446
643.5	.1932	759.2	.7120
645.3	.3418	783.8	.7596
649.6	.3368	817.3	.8321
660.0	.4395	843.6	.8806
664.0	.4500	873.2	.9252
692.7	.5350	905.7	.9725
713.1	.5982		
	54.85%		
607.8	1.2450	754.7	1.6055
612.6	.2585	781.5	.6654
624.2	.2896	807.5	.7141
646.1	.3500	808.7	.7208
658.4	.3834	825.1	.7441
682.1	.4398	839.5	.7715
691.6	.4666	861.0	.8128
707.3	.5004	875.8	.8332
724.6	.5376	903.7	.8815
739.7	.5732		
	74.33%		
634.0	1.2578	779.6	1.5777
644.8	.2837	810.8	.6347
667.8	.3369	829.0	.6547
698.0	.4114	856.9	.7012
722.3	.4683	876.4	.7296
745.8	.5115	904.5	.7631
759.9	.5400		
	84.70%		
668.4	1.3069	772.2	1.5089
686.0	.3493	787.8	.5429
709.1	.3961	794.9	.5517
730.5	.4316	822.9	.6051
737.6	.4439	847.5	.6440
742.2	.4585	870.2	.6772
766.3	.5024	904.0	.7282
	93.96%		
690.5	1.2913	857.0	1.6126
719.8	.3612	877.4	.6388
750.4	.4307	878.4	.6402
769.8	.4701	902.1	.6663
799.1	.5239	902.7	.6666
830.4	.5759		
	100%		
685.4	1.2440	770.1	1.4643
689.0	.2613	783.2	.4955
697.5	.2940	785.7	.4970
698.5	.2961	791.1	.5022
708.6	.3234	815.7	.5440
718.5	.3546	830.7	.5637
738.0	.4041	847.4	.5905
739.5	.4132	877.2	.6232
745.7	.4207	897.6	.6555
754.3	.4295	910.7	.664
768.2	.4639		

Potassium chloride (KCl) + Potassium cyanide (KCN)

Truthe, 1912

%	f.t.	m.t.	%	f.t.	m.t.
0	775	-	65	665	642
10	747	733	80	645	626
25	722	702	90	636	622
38	698	675	100	622	-
50	682	660			

Potassium chloride (KCl) + Potassium hydroxide (KOH)

Scarpa, 1915

wt%	mol%	f.t.	tr.t.	m.t.	E	min.
0	0	776	-	-	-	-
10	12.94	734	-	660	-	-
20	24.65	692	-	560	-	-
30	36.06	645	430	-	125	30
40	47.02	585	432	-	"	80
50	57.05	475	430	410	128	120
60	66.87	430	"	400	130	140
70	75.61	422	-	395	128	170
80	84.02	415	150	390	122	100
90	92.48	402	215	385	120	50
95	96.19	395	245	"	110	20
100	100	380	265	-	-	-

Potassium chloride (KCl) + Potassium metaborate (KBO_2)

Posypaiko ,Bergman and Kislova, 1956.

mol%	f.t.	mol%	f.t.
0	774	35	750
2.5	768	37.5	760
5	760	40	770
7.5	752	45	785
10	745	50	-
12.5	740	55	804
15	735	60	815
17.5	730	65	828
20	726	70	843
22.5	722	75	885
25	715 E	80	865
27	718	85	880
27.5	718	90	900
30	725	100	947
32.5	736		

Potassium chloride (KCl) + Potassium nitrate (KNO_3)

Jänecke, 1923

mol%	f.t.
100	330
90	320
0	790

Jänecke, 1928

mol%	f.t.
100	330
95	320 E
90	345
80	402

Lifshits, 1956.

%	f.t.
100	337
94	320 E
82.6	360° (1+1)
0	774

Potassium chloride (KCl) + Potassium carbonate (K_2CO_3)

Amadori, 1913

%	f.t.	E	min.
100	896	-	-
95	866	626	30
85	804	632	70
80	778	634	90
70	724	635	120
64.96	698	636	140
50	-	636	180
40	660	636	120
25	715	635	80
10	750	632	60
0	774	-	-

Volkov and Bergman, 1940 (fig.)

mol%	f.t.	mol%	f.t.
100	890	17.7	700
53.9	730	0	860
38.0	625		

Potassium chloride (KCl) + Potassium sulfate (K_2SO_4)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	795	30	735
10	760	40	800
20	725	100	1060
24	710		

Jänecke, 1908

mol%	f.t.	E	tr.t.
100	1074	-	587
81.9	1009	700	-
60.0	880	695	590
33.4	715	680	-
26.6	-	691	588
25.0	698	690	590
17.7	710	690	585
11.2	735	690	-
5.3	753	695	-
0	778	-	-

Dombrovskaya, 1933 (fig.)

mol%	f.t.	mol%	f.t.
0	775	42.9	830
11.2	740	66.7	980
26.4	698 E	100	1076

Akopov and Bergman, 1954

mol%	f.t.	mol%	f.t.
100	1069	27.0	692
88.7	1026	23.5	698
78.6	986	20.5	706
69.5	948	18.4	713
60.8	902	17.7	724
53.4	866	13.0	734
46.6	828	10.5	740
40.4	790	8.2	750
35.2	754	5.9	760
32.0	732	3.1	770
29.4	713	0	774

E : 26.6% 690°

Bergman, Kislova and Posipaiko, 1955

mol%	f.t.	mol%	f.t.
100	1069	38	770
94.2	1044	35.2	754
90.5	1026	32.5	732
83.5	1010	29.9	713
78.6	986	27.0	692
74.0	970	24.3	698
69.5	948	20.5	706
65.3	902	18.4	713
61.3	884	15.7	724
56.9	866	13.0	734
53.9	848	10.5	740
50.4	828	8.2	750
47.0	828	6.4	760
43.9	808	3.1	770
40.9	790	0	774

Semenkova and Bergman, 1956

mol %	f.t.	mol %	f.t.
0	770	42.5	694
10	758	45	712
20	738	47.5	728
30	716	49	743 tr.t.
32.5	710	50	744
35	705	55	780
37.5	700	60	816
40	694	65	848
41.5	690 E	66.5	856 tr.t.
-	690 tr.t.	70	884

Bergman and Bakumskaya, 1956

%	f.t.	%	f.t.
0	774	35.2	766
5.9	758	38.0	780
8.2	750	40.4	790
14.3	730	42.9	808
21.2	704	45.5	822
23.1	698	48.2	836
25.0	692	51.0	850
25.9	690 E	52.5	856 tr.t.
27.0	704	53.9	870
31.2	732	56.9	886
33.3	750	60.0	902

Potassium chloride (KCl) + Potassium chromate (K_2CrO_4)

Zhemchuzhni, 1908

mol%	f.t.	E	tr.t.
0	790	-	-
1.78	781.5	656	-
4.19	770.5	656	-
10.3	745	657	-
16.35	720	658	-
22.52	694	-	-
30.00	664	-	-
38.50	682	-	-
46.33	718	-	-
53.52	768	-	676
60.00	804	-	-
70.06	864	-	-
75.50	886	-	-
78.00	906	-	-
80.06	916	-	-
89.80	958	-	-
91.90	960	657	-
100	984	-	679

Sackur, 1912

m	f.t.
0	771
0.61	755
0.96	747
1.24	740

Benrath and Wainoff, 1911

t	$\varkappa.10^6$	t	$\varkappa.10^6$
	0%		
640	5.0	720	36.9
650	6.5	730	43.7
660	8.4	740	59.6
670	12.0	750	77.8
680	13.3	760	91.9
690	18.6	770	133.8
700	20.9	780	199.2
710	28.8	790	1908000

t	$\varkappa.10^6$				
	10%	20%	30%	40%	50%
500	2.4	3.2	5.5	7.0	9.0
510	3.0	4.2	7.2	8.8	11.3
520	3.7	5.1	9.0	11.0	14.3
530	4.5	6.4	11.0	13.5	18.5
540	5.5	7.0	14.5	17.2	23.0
550	7.0	9.5	18.0	21.0	28.0
560	8.3	11.7	21.5	25.5	34.0
570	10.0	14.2	25.5	31.0	40.0
580	12.0	17.5	31.0	37.0	48.0
590	14.5	21.0	37.0	45.0	57.0
600	17.9	28.6	51.0	59.0	68.0
610.5	14.8	32.0	55.0	65.0	83.0
620	22.2	39.0	69.0	79.0	103
630	28.2	50.0	82.0	98.0	132
640	37.0	75.0	103	128	173
650	55.0	130	165	200	231
660	100	5024	10015	6019	7210

t	60%	70%	80%	90%
500	9.0	7.3	6.5	4.5
510	11.0	9.0	7.8	5.8
520	14.0	11.6	9.0	7.0
530	16.8	13.7	10.8	8.7
540	20.0	16.3	11.0	10.2
550	24.0	19.7	15.3	12.2
560	29.0	23.7	18.3	14.5
570	36.0	29.0	22.0	17.0
580	44.0	35.0	26.0	20.0
590	52.0	42.0	31.0	24.0
600	63.0	52.0	37.0	32.0
610.5	77.0	65.0	45.0	35.0
620	96.0	82.0	57.0	44.0
630	122	105	70.0	54.0
640	158	140	88.0	71.0
650	219	200	160	120
660	8012	7200	5000	7000

t	$\varkappa.10^6$	t	$\varkappa.10^6$
	100%		
500	4.5	740	1320
510	4.7	750	1570
520	5.0	760	1800
530	5.4	770	2030
540	6.0	780	2370
550	6.7	790	2680
560	7.6	800	3040
570	8.9	810	3500
580	10.2	820	4000
590	12.1	830	4050
600	14.5	840	5450
605	16.0	850	6400
610	17.8	860	7600
620	22.0	870	9100
630	27.0	880	12000
640	35.0	890	13200
650	45.0	900	16200
660	63.0	910	21000
670	110	920	36000
679	440 (β)	930	45480
690	520	940	53000
700	610	950	63000
710	800	960	94000
720	970	970	200000
730	1150		

Potassium chloride (KCl) + Potassium molybdate (K_2MoO_4)

Bukhalova and Mateiko, 1955

mol%	f.t.	mol%	f.t.
0	775	40.4	650
14.3	718	42.9	670
23.1	684	45.5	682
27.0	673	53.9	740
31.2	654	66.7	807
33.4	644	81.9	876
35.6	630	100	926
38.0	630		

Potassium chloride (KCl) + Potassium tungstate (K_2WO_4)

Bergman, Kislova and Posipaiko, 1955

mol%	f.t.	mol%	f.t.
100	926	31.6	642
90.5	894	29.1	652
81.9	850	27.4	660
74.0	812	25	672
66.7	780	23.5	678
60	746	21.2	686
53.9	712	17.7	704
48.2	678	14.3	716
45	664	11.2	734
42.9	642	8.2	744
40.9	632	5.3	754
38	622	2.6	766
36.1	628	0	774
33.4	638		

Bukhalova and Mateiko, 1956.

mol%	f.t.	mol%	f.t.
100	926	45	685
90.5	898	41.9	664
81.9	872	38.9	646
74.0	836	36.1	621 E
66.7	804	33.4	660
60	772	30.8	670
53.9	740	25.9	680
48.2	700	0	775

Potassium chloride (KCl) + Potassium bichromate ($K_2Cr_2O_7$)

Zhemchuzhni, 1910

mol%	f.t.	E	tr.t.
100	395	-	236
95.68	385	-	235
90.69	378	-	234
84.71	371	-	-
80.00	370	-	-
75.00	368	-	-
70.00	381	366	-
69.05	383	-	-
65.00	410	-	-
58.45	458	-	-
50.00	523	-	-
40.00	582	-	-
30.00	642	-	-
21.30	685	-	-
13.58	722	-	-
6.53	757.5	365	-
2.55	776.5	365	-
0	790	-	-

Potassium bromide (KBr) + Potassium iodide (KI)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	705	70	705
90	"	60	715
80	"	0	750

Amadori and Pampanini, 1911

%	f.t.	%	f.t.
100	680	60	670
95	675	50	676
90	672	40	684
85	668	30	697
80	665	20	710
75	664	10	724
70	664	0	740

Wrzesnewsky, 1911 and 1912

mol%	f.t.	m.t.	pressure of flow after mixing (kg/mm^2)	after 5 weeks
0	748	-	36.5	36.1
10	723	712	66	-
20	692	670	78.5	54.6
30	640	624	82	56.4
40	603	596	80.5	-
45	592	589	-	-
50	589	589	78.2	53.7
55	597	592	-	-
60	607	596	-	-
70	632	612	68.4	-
80	654	636	59.5	-
90	668	660	40	-
100	680	-	21	-

Zhemchuzhni and Rambach, 1910

mol%	Q mix (mixed crystals)
25	3517
50	4070
75	2657
	(after $1^{1/2}$ month)
50	1614

Kurnakov and Zhemchuzhni, 1912

Q mix (equimolar mixed crystals)= 580 cal/mole

Q decomp. (after $1^{1/2}$ h.) = 230 cal/mole

Potassium bromide (KBr) + Potassium hydroxide (KOH)

Scarpa, 1915

wt%	mol%	f.t.	E	min.	tr.t.
0	0	760	-	-	-
2.5	6.94	680	270	20	190
10	19.35	595	295	30	"
20	34.31	520	300	50	195
30	47.74	440	305	80	205
40	58.67	355	"	90	"
50	67.94	310	300	120	195
60	76.43	340	"	100	"
70	83.33	365-325	280	10	"
80	89.37	370-350	-	-	205
90	95.24	375	-	-	230
100	100	380	-	-	265

Potassium bromide (KBr) + Potassium nitrate (KNO_3)

Rostkovski, 1929

mol%	f.t.	mol%	f.t.
100	337	87.7	401
99.3	338	86.8	407
98.6	339	85.1	415
97.3	339	84.3	420
95.3	341	83.4	426
94.9	345	82.7	430
93.9	355	81.3	435
93	363	80	445
92	367	74.5	451
91.2	374	70	474
90.3	380	60	494
89.4	391	50	540
88.5	398		582

Potassium bromide (KBr) + Potassium sulfate (K_2SO_4)

Gromakov, 1951

mol%	f.t.	E	mol%	f.t.	
0	730	730	50	858	-
10	700	670	60	907	-
15	683	670	70	954	-
20	670	670	80	995	-
30	740	670	90	1038	-
40	804	670	100	1074	1074

Potassium bromide (KBr) + Potassium chromate (K_2CrO_4)

Gromakov, 1951

mol%	f.t.	E	mol%	f.t.	E
0	730	730	34	660	620
5	720	620	38	683	620
10	704	620	50	745	620
14	691	620	60	800	620
18	674	620	70	846	620
22	660	620	(80)	895	-
26	644	620	(90)	935	-
(29)	(620)	-	100	975	975
30	628	620			

Potassium iodide (KI) + Potassium hydroxide (KOH)

Scarpa, 1915

wt%	mol%	f.t.	E	min.	tr.t.
0	0	695	-	-	-
2.5	6.94	665	245	20	-
10	25.00	580	250	45	-
20	43.17	485	255	60	-
30	55.79	365	255	90	-
40	66.45	290	250	130	-
50	74.79	270	245	140	260
60	81.68	310	"	100	"
70	87.41	335	"	60	"
80	92.25	360	"	40	"
90	96.38	375	"	20	265
100	100	380	-	-	"

Potassium iodide (KI) + Potassium sulfate (K_2SO_4)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	705	20	705
10	670	30	770
13	635	100	1060

Gromakov, 1951

mol%	f.t.	E	mol%	f.t.	E
0	680	680	50	854	640
5	662	640	(60)	946	-
10	648	640	(70)	975	-
(12)	640	-	(80)	1006	-
15	676	640	(90)	1040	-
20	730	640	100	1070	-
40	800	640			

Nyankovskaya, 1956.

mol%	f.t.	mol%	f.t.
0	678	25	656
3	677	30	706
5	675	35	742
6	672	40	775
9	670	45	803
10	669	50	825
12	667	55	843
15	664	60	860
20	656	100	1076
24	648		

Potassium iodide (KI) + Potassium chromate (K_2CrO_4)

Gromakov, 1951

mol%	f.t.	E	mol%	f.t.	E
0	680	680	35	680	600
5	660	600	40	717	600
10	640	600	50	768	-
15	626	600	60	819	-
20	608	600	70	860	-
(22)	600	-	(80)	900	-
25	616	600	(90)	940	-
30	650	600	100	975	-

Potassium thiocyanate (KCNS) + Potassium nitrite (KNO_2)

Palkin, 1945

mol %	f.t.	mol %	f.t.
0	178	40	128
5	173	45	155
10	164	48	178
15	157	50	194
17.8	153	55	224
20	148	60	253
25	139	65	279
30	129	85	370-380 (explosion)
35	124	100	422

E : 34 % 120° (1+1)

Potassium sulfide (K_2S) + Potassium sulfate (K_2SO_4)

Goubeau, Kolb and Kroll, 1938

mol%	f.t.	m.t.	E	min.
100	1067	-	-	-
95.5	1047	1025	-	-
80.4	1011	-	-	-
79.5	967	847	-	-
69.0	916	750	-	-
61.9	872	610	-	-
49.8	799	-	587	20
45.8	727	-	"	50
39.6	689	-	"	"
31.2	647	-	"	80
29.9	684	-	586	40
27.2	727	-	"	30
25.8	755	-	587	20
22.8	759	-	-	00
19.7	725	-	725	40
19.3	727	-	727	50
18.4	747	-	"	30
16.2	770	-	729	50
14.3	790	-	737	40
12.9	810	-	725	40
11.4	827	-	727	20
10.0	831	-	737	30
2.2	895	-	-	00
0	912	-	-	-

Next Page

Potassium hydroxide (KOH) + Potassium nitrate (KNO_3)

Retortillo and Moles, 1933

mol%	f.t.	mol%	f.t.
100	343	48.4	225
92.2	300	47.08	222
89.8	278	42.2	227.5
89.1	290	42.08	226.5
88.5	267.5	42.0	228
87.4	268	38.3	231
82.5	198.3-203	34.9	229.8
82.1	210	34.8	228
78.7	207	34.5	229
74.8	218-197	33.0	231.5
70.9	208	32.89	228.5
70.7	207	32.7	228
67.6	209.7	31.4	221-228
64.5	213	31.2	227.5
56.4	218-216.5	28.9	227.5
55.3	216	21.9	200-210
49.5	230-226	18.7	209
49.2	221.5	12.2	250
50.0	218	10.5	262.5
48.8	214	0	336

(2+1)

Potassium hydroxide (KOH) + Potassium carbonate (K_2CO_3)

Seward and Martin, 1949

%	f.t.	E
1.0	408.0	358.7
4.9	400.0	365.0
7.0	394.5	365.0
12.0	384.2	366.9
16.2	373.5	366.5

Potassium metaborate (KBO_2) + Potassium metaphosphate (KPO_3)

van Klooster, 1910

%	f.t.	E	%	f.t.	E
100	810	-	50	871	-
95	752	677	45	872	-
90	-	681	40	850	-
85	708	676	35	800	727
80	-	-	30	-	770
75	-	-	25	844	766
70	-	-	20	872	719
65	848	-	15	891	705
60	885	-	10	915	677
55	868	-	0	947	-

(1+1)

Potassium metaborate (KBO_2) + Potassium sulfate (K_2SO_4)

Kislova, Posipaiko and Bergman, 1955

mol%	f.t.	mol%	f.t.
0	947	25.0	852
1.3	930	27.0	848
2.6	924	29.1	858
3.9	918	31.2	864
5.3	910	33.4	872
6.7	904	35.6	878
8.2	898	38	884
9.6	892	40.4	890
11.2	886	42.9	900
12.6	882	48.2	918
14.3	878	53.9	936
16.0	875	60.0	956
17.7	872	66.7	978
19.5	866	91.9	1000
21.2	862	100	1069
23.1	857		

Potassium metaborate (KBO_2) + Potassium tungstate (K_2WO_4)

Bergman, Kislova and Posipaiko, 1955

mol%	f.t.	mol%	f.t.
100	926	33.4	840
95.1	910	31.2	844
90.5	898	29.1	848
86.1	890	27	852
81.9	878	25	856
77.8	862	23.1	860
74	862	21.2	864
70.2	857	19.5	868
66.7	852	17.7	872
63.7	846	16.0	880
60	840	14.3	884
56.8	834	12.6	896
53.9	830	11.2	894
50.5	826	9.6	900
48.2	820	8.2	905
45.5	816	6.7	910
42.9	808	5.3	916
40.4	820	3.9	920
38	826	2.6	926
35.6	834	1.3	936
		0	947

Potassium nitrite (KNO_2) + Potassium nitrate (KNO_3)

Meneghini, 1912

%	f.t.	m.t.	%	f.t.	m.t.
10.97	424	-	82.5	320.5	-
17.5	413	-	85	320	90
26.1	398.5	-	88.5	320	91.5
35	383.5	-	93.5	324	98
55.5	354	-	97.5	330	110
73.9	328	-	100	336	120
80.3	321.5	-			

Previous Page

Potassium nitrate (KNO_3) + Potassium carbonate (K_2CO_3)

Amadori, 1913

%	f.t.	E	min.	tr.t.
0	336	-	-	124
5	-	326	300	124
10	392	326	280	122
15	454	326	220	122
20	496	325	180	122
25	552	325	150	122
35	622	320	140	120
40	655	316	120	120
50	700	312	100	116
75	-	306	70	
100	896	-		

Potassium nitrate (KNO_3) + Potassium sulfate (K_2SO_4)

Amadori, 1913

%	f.t.	E	min.
0	336	-	-
2	334	-	-
5	-	332	320
7.5	382	"	280
10	448	"	280
15	544	"	250
20	594	330	220
30	666	"	180
50	-	328	130
65	-	326	100
100	1066	322	-

Potassium nitrate (KNO_3) + Potassium tungstate (K_2WO_4)

Bergman, Kislo\a and Posipaiko, 1955

mol%	f.t.	mol%	f.t.
0	337	8.2	370
1.1	332	9.3	392
2.6	326	11.2	416
3.7	321	12.4	438
5.3	318	14.3	468
6.4	328	17.7	514
7.6	350	19.1	decomposition

Potassium metaphosphate (KPO_3) + Potassium pyrophosphate ($K_4P_2O_7$)

Parravano and Calcagni, 1909

%	f.t.	E	min.
0	823	-	-
2	798	-	-
4	784	-	-
10	778	-	-
12	771	-	-
21	731	-	-
23	718	-	-
25	708	-	-
26	702	-	-
28	694	-	-
30	683	580	160
36	655	603	260
38	643	615	320
41	-	"	400
44	-	"	500
46	-	"	460
50	-	"	380
52	680	"	340
54	703	618	300
56	717	620	260
58	728	615	240
60	752	615	220
64	739	620	180
74	898	614	60
80	951	610	20
85	993	-	-
90	1034	-	-
100	1092	-	-

Morey, 1954

%	f.t.	
0	813	
20.14	754	
41	613	E
41.98	625 (1+1)	tr.t.
46	641.5	
46.00	644	
50.23	702	
73.66	940	
82.00	995	
100	1109	

Potassium metaphosphate (KPO_3) + Potassium sulfate (K_2SO_4)

Bergman and Sholokhovich, 1953

mol %	f.t.	mol %	f.t.
0	798	17.3	764
1.1	790	21.2	790
2.1	782	25	820
3.1	776	29.1	852
5.3	766	33.4	874
6	758	37	888
8.7	752	42.9	908
11.2	736	47.1	926
11.8	728	53.9	941
13	728	60	961
15	744	100	1069

E : 12.4 % 718°

Potassium niobate ($KNbO_3$) + Potassium tantalate ($KTaO_3$)

Reisman, Triebwasser and Holtzberg, 1955

mol %	f.t.	m.t.
0	1039	1039
20	1136	1070
30	-	1087
35	1194	-
40	-	1115
50	1243	1141
65	1282	-
75	-	1223
80	1318	1247
100	1357	1357

Potassium metavanadate (KVO_3) + Potassium metatitanate (K_2TiO_3)

Sholokhovich and Barkova, 1956 (fig.)

mol%	f.t.	mol%	f.t.
100	826	7.0	850
96.1	788 E	1.3	512 E
66.7	820 tr.t.	0	520
33.3	1006		
		(1+1)	

Potassium chlorate ($KClO_3$) + Potassium nitrate (KNO_3)

Herbette, 1906

%	angles (001) - (T11)	angles (T11) - (11T)
0	74.25	75.38
21	74.11	76.7
25.75	73.50	76.10
27	73.39	76.27
35	73.25	77.30
100	73.30	73.30

Potassium perchlorate ($KClO_4$) + Potassium permanganate ($KMnO_4$)

Sommerfeldt, 1899 - 1901

Q mixed crystals = 0

Potassium pyrophosphate ($K_4P_2O_7$) + Potassium molybdate (K_2MoO_4)

Belyaev and Sholokhovich, 1953

mol%	f.t.	mol%	f.t.
100	926	70.9	866
97.4	911	66.6	883
94.7	898	62	910
91.8	888	57.1	932
88.8	870	51.8	956
85.7	854	48.4	968
82.3	842	46.1	982
78.8	825	45.3	978
76.5	830	40	995
75	845	33.3	1018

E : 78% 822°

Potassium pyrophosphate ($K_4P_2O_7$) + Potassium tungstate (K_2WO_4)

Sholokhovich and Bergman, 1954 (fig.)

mol%	f.t.	mol%	f.t.
100	928	66.6	928
85.7	866	40	1015
78.8	836	0	1095

Potassium carbonate (K_2CO_3) + Potassium sulfate (K_2SO_4)

Le Chatelier, 1894 and 1897

mol%	f.t.	mol%	f.t.
0	860	67	960
33	880	75	980
40	900	100	1045
50	920		

Amadori, 1913

%	f.t.	tr.t.	%	f.t.	tr.t.
0	896	415	75	995	628
10	902	-	80	1010	622
20	907	-	85	1024	608
34	920	-	88	1033	598
40	926	-	90	1041	595
46	935	-	92	1045	593
50	938	-	95	1052	586
55.77	945	-	97.5	1060	585
60	964	-	100	1066	583
68	975	-			

Potassium metatitanate (K_2TiO_3) + Potassium sulfate (K_2SO_4)

Sholokhovich and Barkova, 1956 (fig.)

mol %	f.t.
0	826
12.0	796
55.0-99.0	917 C+L_1+L_2
100	966

Potassium metatitanate (K_2TiO_3) + Potassium chromate (K_2CrO_4)

Sholokhovich and Barkova, 1956 (fig.)

mol %	f.t.
0	826
12.0	796
55.0-99.0	917 C + L_1 + L_2
100	966

Potassium metatitanate (K_2TiO_3) + Potassium molybdate (K_2MoO_4)

Sholokhovich and Barkova, 1956 (fig.)

mol %	f.t.
0	826
17	786 E
40.0-99.0	917 C + L_1 + L_2
100	926

Potassium metatitanate (K_2TiO_3) + Potassium tungstate (K_2WO_4)

Sholokhovich and Barkova, 1956 (fig.)

mol %	f.t.
0	826
19.0	787
25.0-94.0	884 C + L_1 + L_2
100	928

Potassium sulfate (K_2SO_4) + Potassium fluoberyllate (K_2BeF_4)

Levina, Novoselova and al., 1956

mol%	f.t.	m.t.	tr.t. 1	tr.t. 2	tr.t. 3
100	790	-	685	-	-
95	786	-	483	-	668
90	808	761	-	-	665
85	819	-	-	-	668
80	840	-	-	-	670
75	858	-	-	-	668
70	862	-	-	-	653
65	870	-	-	-	649
60	880	-	-	-	645
55	886	-	-	-	647
50	916	875	-	812	637
45	938	-	-	794	621
40	942	-	-	804	613
35	955	-	-	-	612
30	976	-	-	804	605
25	977	-	-	805	601
20	1002	-	-	825	596
15	1009	-	-	-	585
10	1033	-	-	854	592
5	1058	1026	-	854	580
3.	-	1041	-	-	-
1	1065	-	854	854	-
0	1071	-	585	-	-

Potassium sulfate (K_2SO_4) + Potassium chromate (K_2CrO_4)

Le Chatelier, 1897

mol%	f.t.	mol%	f.t.
100	940	50	985
85	950	34	1003
67	960	0	1045

Groschuff, 1908

%	f.t.	m.t.	tr.t.
100	971	-	666
80	975	965	652-644
60	996	989	640-637
40	1017	1011	625-624
20	1038	1032	605
0	1072	-	586

Amadori, 1913

%	f.t.	tr.t.	%	f.t.	tr.t.
100	978	666	40	1026	623
90	981	661	20	1045	600
80	984	651	10	1056	592
60	1002	639	0	1066	585
50	1016	632			

Stibing, 1906

		crystalline axes		
%	d	a	b	c
0	-	0.5727	1	0.7418
0.51	2.6664	.5727	-	.7436
1.21	.6669	.5723	-	.7445
2.71	.6680	.5719	-	.7444
9.39	.6732	.5715	-	.7418
40.54	.6959	.5712	-	.7381
51.55	.7345	.5704	-	.7351
100	-	.5696	-	.7351

Potassium sulfate (K_2SO_4) + Potassium molybdate (K_2MoO_4)

Amadori, 1913

%	f.t.	tr.t.	%	f.t.	tr.t.
100	926	475	60	934	484
95	922	460	50	952	508
90	920	-	40	980	525
80	"	-	30	998	544
75	"	-	20	1018	558
70	922	465	10	1042	568
65	926	475	0	1066	583

Bergman, Kislova and Korobka, 1954

mol%	f.t.	mol%	f.t.
100	924	40	954
90	912	30	978
80	906	25	993
75	903	20	1002
65	909	10	1038
55	924	0	1069
50	933		

Potassium sulfate (K_2SO_4) + Potassium tungstate (K_2WO_4)

Amadori, 1913

%	f.t.	tr.t.	%	f.t.	tr.t.
100	894	575	55	935	504
95	890	545	50	947	510
92	886	525	35	984	528
88	884	-	30	994	545
85	884	-	20	1015	558
75	892	-	10	1045	572
70	904	-	0	1066	583
65	912	482			

Bergman, Kislova and Posipaiko, 1955

mol%	f.t.	mol%	f.t.
100	926	70.0	916
97.5	924	67.5	916
95.0	922	65.0	920
92.5	922	62.5	924
90.0	922	59.5	924
87.5	922	57.5	928
85.0	920	55.0	934
82.0	920	50.0	944
80.0	920	45.0	956
77.5	918	35.0	980
75.0	918	0	1069
72.5	916		

Potassium chromate (K_2CrO_4) + Potassium bichromate ($K_2Cr_2O_7$)

Groschuff, 1908

%	f.t.	tr.t.	min.	E	min.
0	971	666	110	-	-
25	909	665	80	390	175
50	812	665	40	393	280
75	675	665	10	393	400
90	504	-	-	394	450
95	444	-	-	392	460
97.5	414	-	-	393	495
98.75	402	-	-	394	495
100	396	-	-	-	-

Amadori, 1913

%	f.t.	tr.t.	E
100	398	-	-
97.5	-	-	394
95	437	-	"
90	503	-	392
80	619	-	393
70	696	-	392
60	758	664	"
40	858	663	"
20	923	665	390
0	978	666	-

Potassium chromate (K_2CrO_4) + Potassium molybdate (K_2MoO_4)

Amadori, 1913

%	f.t.	tr.t.	%	f.t.	tr.t.
100	926	475	50	939	540
95	"	465	40	942	575
90	"	-	30	955	600
80	"	-	20	960	624
75	"	-	10	972	648
70	930	470	0	978	666
60	934	502			

Potassium chromate (K_2CrO_4) + Potassium tungstate (K_2WO_4)

Amadori, 1913

%	f.t.	tr.t.	%	f.t.	tr.t.
100	894	595	65	910	516
94	"	530	50	922	555
90	"	-	40	938	585
88	"	-	25	960	622
80	899	-	15	966	640
75	902	-	0	978	666
70	906	494			

Kislova, Posipaiko and Bergman, 1955

mol%	f.t.	mol%	f.t.
100	926	60.0	936
90.0	928	50.0	940
80.0	928	40.0	942
70.0	930		

Potassium molybdate (K_2MoO_4) + Potassium tungstate (K_2WO_4)

Amadori, 1913

%	f.t.	tr.t.	%	f.t.	tr.t.
100	894	575	40	916	-
90	896	560	30	920	-
80	899	540	20	920	475
70	902	-	10	924	475
60	906	-	0	926	475
50	912	-			

Potassium pyrosulfate ($K_2S_2O_7$) + Potassium acid sulfate ($KHSO_4$)

Cambi and Bozza, 1923

%	f.t.	tr.t. I	tr.t. II	tr.t. III
0	414.2	315	226	-
3	401.7	300	224	-
8.51	380.0	310	-	193.6
22.34	328.0	-	-	"
31.87	308.2	-	-	185
37.95	299.7	-	-	201
47.34	285.0	-	213	194.5
56.35	264.5	-	-	198
61.66	254.0	-	-	200.5
67.12	242.4	-	-	201.2
68.68	236.5	-	-	201.7
64.91	223.6	-	-	199.8
76.11	224.0	-	-	199.0
78.25	227.3	-	-	200.5
79.47	220.2	-	-	199.3
81.70	217.7	-	-	198.0
83.68	212.6	-	-	201.2
86.30	202.6	-	-	201.2
88.30	201.3	-	-	198.6
91.40	211.5	-	-	201.2
95.34	212.0	-	-	-
98.46	213.2	-	-	-
100	213.8	-	-	-

Potassium bichromate ($K_2Cr_2O_7$)
+ Potassium bimolybdate ($K_2Mo_2O_7$)

Amadori, 1913

%	f.t.	%	f.t.
0	398	60	440
10	398	70	450
20	408	80	458
30	418	90	472
40	425	100	484
50	433		

Potassium bichromate ($K_2Cr_2O_7$)
+ Potassium bitungstate ($K_2W_2O_7$)

Amadori, 1913

mol%	f.t.	mol%	f.t.
0	398	34.54	412
5.54	392	44.22	434
11.65	388	55.18	454
14.95	384	67.85	492
18.44	384	82.60	512
22.12	386	100	555
26.02	392		

Potassium bimolybdate ($K_2Mo_2O_7$)
+ Potassium bitungstate ($K_2W_2O_7$)

Amadori, 1913

mol%	f.t.	mol%	f.t.
0	484	61.75	524
7.14	490	79.68	537
10.88	492	86.16	542
22.87	502	100	555
40.90	510		

Potassium acetate ($KC_2H_3O_2$)
+ Potassium butyrate ($KC_4H_7O_2$)

De Mol, 1925 (fig.)

mol%	f.t.	E	mol%	f.t.	E
0	306	-	30	302.5	277.5
10	289	-	40	314	-
17	277.5	277.5	46	318	-
20	282	277.5			

Potassium caproate ($KC_6H_{11}O_2$)
+ Potassium laurate ($KC_{12}H_{23}O_2$)

Hesz and Kiessig, 1948 (fig.)

Mixed crystals

mol%	lattice plane distance
100	31
80	29
60	26
40	24
20	21
0	19

Potassium caproate ($KC_6H_{11}O_2$)
+ Potassium oleate ($KC_{18}H_{33}O_2$)

Hesz and Kiessig, 1948 (fig.)

Mixed crystals

mol%	lattice plane distance
100	44.0
80	39
60	33
40	29
20	23
0	19.0

Potassium laurate ($KC_{12}H_{23}O_2$)
+ Potassium oleate ($KC_{18}H_{33}O_2$)

Hesz and Kiessig, 1948 (fig.)

Mixed crystals

mol%	lattice plane distance
100	44.0
80	41.0
60	39.0
40	37
20	34
0	31.0

Rubidium fluoride (RbF) + Rubidium carbonate (Rb_2CO_3)

Schmitz-Dumont and Heckmann, 1949 (fig.)

mol %	f.t.	mol %	f.t.
100	850	40	630 (1+1)
80	785	34	616 E"
60	675	20	700
51	627 E	0	795
50	636		

Rubidium fluoride (RbF) + Rubidium sulfate (Rb_2SO_4)

Schmitz-Dumont and Heckmann, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	1055	40	845
80	965	20	770
60	870	14	723
57	847	0	800
50	856 (1+1)		

Rubidium fluoride (RbF) + Rubidium chromate (Rb_2CrO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	994	-	50	783	-
80	900	775	40	770	704
60	795	"	20	704	"
56	775	"	0	785	-

tr.t. B = 552° tr.t.(1+1) = 637°

Rubidium fluoride (RbF) + Rubidium molybdate (Rb_2MoO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	958	-	40	750	690
80	870	756	23	690	"
59	756	"	10	750	"
50	762	-	0	790	-

tr.t. B = 552°

Rubidium fluoride (RbF) + Rubidium tungstate (Rb_2WO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	961	-	40	760	729
80	870	769	20	729	"
57	769	"	0	800	-
50	772	-			

Rubidium fluoride (RbF) + Rubidium bititanate ($Rb_2Ti_2O_5$)

Schmitz-Dumont and Schulz, 1952 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	870	-	30	820	780
80	840	705	20	800	780
71	820	820	5	780	780
60	840	-	0	798	-
50	835	740	(2+3)		

Rubidium chloride (RbCl) + Rubidium nitrate ($RbNO_3$)

Rostovskii, 1930

%	f.t.	tr.t. I	tr.t. II	tr.t. III
100	317	316.5	290	222
98.3	316	-	-	229
96.5	324	-	-	220
94.8	330	-	-	-
92.2	333	-	-	-
89.6	338	-	-	-
86.8	354	-	-	-
83.7	365	-	-	-
80.8	376	340	-	183
76.1	394	-	-	-
75	410	-	-	-
72.5	430	-	-	-
69.3	443	-	-	-
66.4	460	325(32.3%)	-	180
63.8	475	-	-	-
61.5	486	-	-	-
59.1	500	-	-	-
56.8	510	-	-	-
54.5	522	394	-	176
50	540	390	-	174

Cesium fluoride (CsF) + Cesium carbonate (Cs_2CO_3)

Schmitz-Dumont and Heckmann, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	800	43	530 E
80	710	20	640
60	615	0	705

Cesium fluoride (CsF) + Cesium sulfate (Cs_2SO_4)

Schmitz-Dumont and Heckmann, 1949 (fig.)

mol%	f.t.	mol%	f.t.
100	1020	40	790
80	925	20	675
60	820	16	648
54	793	0	705
50	798 (1+1)		

Cesium fluoride (CsF) + Cesium chromate (Cs_2CrO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	956	-	40	740	641
80	895	742	30	710	"
60	775	"	19	641	"
56	742	"	0	705	-
50	747	-	(1+1)		

Cesium fluoride (CsF) + Cesium molybdate (Cs_2MoO_4)

Schmitz-Dumont and Weeg, 1951 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	936	-	40	735	635
80	880	740	20	660	"
60	780	"	15	635	"
53	740	"	0	700	-
50	742	"			

tr.t. 1 = 547 2 = 425 (1+1)

Cesium fluoride (CsF) + Cesium tungstate (Cs_2WO_4)

Schmitz-Dumont and Weeg, 1951

mol%	f.t.	E	mol%	f.t.	E
100	953	-	40	748	639
80	865	752	20	660	"
60	765	"	16	639	"
57	752	"	0	705	-
50	755	"			

tr.t. B = 542 tr.t. (1+1) = 461

Cesium fluoride (CsF) + Cesium bititanate ($Cs_2Ti_2O_6$)

Schmitz-Dumont and Schulz, 1952 (fig.)

mol%	f.t.	E	mol%	f.t.	E
100	895	-	30	735	600
75	850	825	20	610	610
65	825	825	10	675	560
60	845	-	0	680	-
50	815	600			

Cesium chloride (CsCl) + Cesium bromide (CsBr)

Wood and Frank, 1951

Mixed crystals

%	n_D^{25} after heating at 400° during 0.5 h	4h	24h
0	1.644	1.657	1.661
10	.6480	.664	.663
20	.6481	.668	.6675
30	.6482	.670	.6676
40	.667	.670	.6678
50	.674	.6701	.6679
60	.676	.6702	.6682
70	.677	.6703	.669
80	.678	.671	.670
90	.6785	.673	.671
100	.689	.679	.673

Cesium chloride (CsCl) + Cesium sulfate (Cs_2SO_4)

Dergunov, 1951

mol%	f.t.	mol%	f.t.
0	636	38.0	688
2.6	628	42.9	712
5.3	620	45.5	724
8.2	605	48.2	736
11.2	588	53.9	758
14.3	567	56.9	784
16.3	553	60	812
17	553	66.7	867
17.7	560	74.0	918
21.2	590	81.9	965
25	618	90.5	998
29.1	642	100	1020
33.4			

Thallium chloride (TlCl) + Thallium bromide (TlBr)

Mönkemeyer, 1906

%	f.t.	%	f.t.
0	426	64.01	421
11.64	422	73.44	426
22.84	419	82.58	433
33.70	416	91.43	442
44.15	413	100	450
54.25	416		

Favorski , 1941 (fig.)

mol%	f.t.	mol%	f.t.
0	428	60	423 E
20	423	80	440
40	423	100	456

Thallium chloride (TlCl) + Thallium iodide (TlI)

Mönkemeyer, 1906

%	f.t.	E	min.
0	426	-	-
13.31	400	316	30
25.67	379	316	50
37.19	354	315	75
47.94	334	316	105
58.01	-	315	140
67.45	338	316	90
76.32	362	316	55
84.68	384	-	-
92.56	410	-	-
100	431	-	-

Barth and Lunde,

Lattice constant study, the lattice constant of mixture is higher than of pure TlI

Thallium chloride (TlCl) + Thallium sulfate (Tl_2SO_4)

Sementsova, Bergman and Lesnykh, 1956.

mol%	f.t.	mol%	f.t.
0.0	430	29.1	394
5.3	426	33.3	412
11.2	412	38.0	428
17.7	390	48.2	468
19.5	380	60.0	506
21.2	368	74.0	550
22.7	358 E	90.5	598
23.5	362	100.0	638

Thallium bromide (TlBr) + Thallium iodide (TlI)

Mönkemeyer, 1906

%	f.t.	%	f.t.
0	450	63.60	411
11.46	442	73.10	412
22.56	430	82.22	420
33.30	420	91.29	424
43.72	413	100	431
53.81	412		

Barth and Lunde, 1926 (fig.)

mol%	lattice constant (in Å)
0	3.96
25	4.04
50	4.09
75	4.16
92	4.19
100	4.18

Thallium bromide (TlBr) + Thallium nitrate ($TlNO_3$)

Rostkovski, 1929

mol%	f.t.	mol%	f.t.
100	206	74.2	268
98.8	205	71.8	277
97.7	202	69.4	286
94.5	200	67.2	294
92	197	63.2	309
89.3	195	59.7	322
88	193	56.5	332
86.3	205	53.7	340
84.6	221	50	350
82.9	229	47.7	357
81.3	240	42.1	375
79.8	246	31.3	395
78.3	253	30.2	404
76.9	258		

Thallium iodide (TlI) + Thallium nitrate ($TlNO_3$)

van Eyk, 1900-1

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	206	205.4	64.0	321.5	215.5
98.4	207	206	50	346.5	215.5
95.9	208	207	45.5	354	-
93.3	211	208.4	39.8	363	-
90.1	215	209.5	30.1	378	290
87.0	238	210	19.5	393	335
82.0	264	215.5	9.9	408	376
76.3	288	215.5	5.3	415	-
68.6	311	215.5	0	422	-

mol% L	mol% C	mol% L	mol% C
95.3	92	86.7	37
93.3	89	80.0	33
90.8	84	75.3	23

Thallium nitrite ($TlNO_2$) + Thallium nitrate ($TlNO_3$)

Cuttica, 1920

%	f.t.	m.t.	tr.t. beginn.	tr.t. end
100	206	205	143.5	143
95	205	197	112	106
93.5	-	-	105	95
90	202.5	191	-	-
85	201	189	-	-
80	199	184	-	-
75	197	182	-	-
70	195.5	179	-	-
65	193.5	177	-	-
60	191.5	174	-	-
55	187	170	-	-
50	185.5	168	-	-
45	180	164	-	-
40	175	160	-	-
35	172	157	-	-
30	166	155	-	-
25	164	151	-	-
20	161	149	-	-
15	158	147	113	107
10	156	145	118	112
5	153	148	122	118
0	149	149	131	130

Thallium nitrate ($TlNO_3$) + Thallium carbonate (Tl_2CO_3)

Brown, 1933

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	206.5	-	50	214	-
1.5	212	-	53.9	213	-
3.6	214	-	57	212	211
5.5	214	-	60.3	213	-
7.8	212	201	63.6	215	211
10	211	-	66.7	222	-
12.7	205	201	68	224	-
15.6	204	201	70.5	229	-
19	208	201	73	233	-
22.8	212	201	75	236	211
26.4	215	-	78.7	242	211
30	217	-	81.3	246	-
34	218	-	84.3	249	-
38	213	-	90.2	256	211
42.1	212	210	100	272	211
46.5	213	-	(2+1)		

Cuprous chloride (CuCl) + Cuprous bromide (CuBr)

Mönkemeyer, 1906

%	f.t.	%	f.t.	tr.t.
0	419	68.49	424	-
13.87	412	77.18	436	373
26.60	409	85.29	450	378
38.32	408	92.88	466	381
49.49	410	100	480	384
59.18	414			

Cuprous chloride (CuCl) + Cuprous iodide (CuI)

Mönkemeyer, 1906

%	f.t.	E	min.	tr.t.
0	419	-	-	-
17.60	384	284	20	214
32.46	353	283	35	"
45.18	320	284	60	"
56.17	320	284	75	"
65.79	372	284	75	"
74.25	424	284	60	"
81.77	482	284	40	216
88.49	533	284	20	238
94.54	590	-	-	372
100	590	-	-	400

Cuprous chloride (CuCl) + Cuprous sulfide (Cu_2S)

Truthe, 1912

%	f.t.	E	min.
0	423	-	-
5	414	389	8.9
10	404	388	18.0
15	392	390	33.3
17	391	391	35.0
25	524	391	24.7
35	637	392	15.8
50	-	393	14.4
55	817	393	12.1
60	-	391	10.2
70	919	392	5.5
80	996-956	393	2.3
90	1057-996	381	0.9
100	1114	-	-

Urazov and Chelidze, 1941

%	f.t.	E	min.
0	422	-	-
10	404	389	9
17	-	387	16
20	459	389	15
30	577	389	13
35	624	387	12
40	641	389	11
50	767	387	9
60	849	386	8
70	-	387	6
100	1150	-	-

Cuprous chloride (CuCl) + Cuprous oxide (Cu_2O)

Truthe, 1912

%	f.t.	E	min.
0	423	-	-
5	-	424	30.5
15	609	423	24.0
25	738	424	23.0
30	780	425	18.8
45	963	423	15.7
55	1025	422	5.5
70	1143	424	3.7
85	1192	423	1.2
93	1215	423	0.6
100	1230	-	-

Cuprous bromide (CuBr) + Cuprous iodide (CuI)

Mönkemeyer, 1906

%	f.t.	tr.t.	%	f.t.	tr.t.
0	480	384	66.55	480	357
12.85	460	374	75.58	502	362
24.90	449	368	84.14	530	373
36.25	443	364	92.27	560	384
46.93	448	360	100	590	400
57.02	460	358			

Silver chloride (AgCl) + Silver bromide (AgBr)

Mönkemeyer, 1906

%	f.t.	%	f.t.
0	452	66.28	413
12.99	444	75.36	413
24.68	435	83.98	415
35.96	428	92.18	419
46.70	420	100	422
56.72	416		

Boardman, Dorman and Heyman, 1949

mol%	molar volume (in cc)
	600°
0	30.50
22.7	31.50
40.3	32.20
65.8	33.30
100	34.75

Kurnakov and Zhemchuzhni, 1908

mol%	pressure of flow Kg/mm	hardness
0.0	17.20	34.0
10.0	28.20	-
25.0	29.00	41.0
50.0	31.86	42.4
75.0	27.97	37.0
90.0	22.03	-
100.0	18.64	30.4

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ	
	500°	600°
0	177	171
20	170	165
40	164	160
60	159	156
80	156	152

Harrap and Heyman, 1955

mol%	η	κ (mhos)	η	κ (mhos)
	440°		480°	
0	-	-	217.6	3.82
20.0	-	3.48	-	3.63
26.8	2580	-	232	-
40.0	2750	3.30	245	3.43
60.0	2970	3.13	263	3.26
80.0	3120	2.96	275	3.11
100	3380	2.87	300	2.98
	520°		560°	
0	1980	3.98	181	4.10
20.0	-	3.77	-	3.89
26.8	2100	-	192	-
40.0	2180	3.56	194	3.66
60.0	2350	3.36	212	3.46
80.0	2460	3.21	224	3.30
100	2690	3.07	245	3.15
	600°			
0	1680	4.21		
20.0	-	3.99		
26.8	1740	-		
40.0	1740	3.77		
60.0	1980	3.55		
80.0	2090	3.38		
100	228	3.21		

Sandonnini, 1913 and 1920

%	κ (melt)(mhos)
500°	
100	2.924
75.35	3.130
57.0	3.246
36.0	3.409
0	3.653

mol%	κ.10^5 (solid)				
	200°	250°	300°	350°	400°
0	5	20	109	520	2400
10	8	35	168	549	3180
30	13	66	350	1680	5700
40	16	92	552	2200	7530
50	20	126	701	3200	12000
60	23	156	817	3700	14000
70	32	210	1080	4250	16400
80	38	231	1190	5200	23500
90	42	240	1290	5850	27200
100	46	260	1460	6420	30400

Silver chloride (AgCl) + Silver iodide (AgI)

Mönkemeyer, 1906

%	f.t.	E	min.	tr.t.
0	452	-	-	-
15.77	419	211	20	112
28.60	381	"	30	111
41.50	343	"	35	112
52.22	293	"	40	113
62.08	251	"	50	115
71.07	251	"	55	115
75.26	371	"	35	115
86.74	435	"	20	115
93.34	495	-	-	120
100	552	-	-	143

Tubandt and Lorenz, 1914

mol%	f.t.	m.t.	tr.t.
0	451		
10	420	211	-
20	383	"	-
30	343	"	-
40	302	"	-
50	256	"	-
58	211	"	-
60	240	"	-
70	350	"	-
80	435	"	-
88	480	112	115
90	495	356	120
100	552	-	143

Bergmann and Gönke, 1926

mol%	f.t.	E	tr.t.
0	451	-	-
10	418	264	2
25	368	263	128
40	318	264	121
50	277	"	130
60	310	263	130
75	404	264	129
96	532	-	138
100	556	-	148

Tubandt and Lorenz, 1914

t	κ(mhos)	t	κ(mhos)
	100%		
600	4.16	400	0.026
550	4.05	350	0.0065
500	3.91	300	0.0015
456	3.76	250	0.00030
450	0.11		

t	κ (mhos)			
	90	75	58	10
	mol%			
600	2.48	2.52	2.85	3.83
550	.42	.47	.81	.70
500	-	-	.75	.54
498	2.35	-	-	-
490	.40	-	-	-
480	.43	-	-	-
458	.45	-	-	-
450	.43	2.34	2.68	4.38
425	-	-	-	3.28
402	-	2.25	-	-
400	2.28	-	2.56	1.15
390	-	2.25	-	-
380	-	-	-	-
365	-	2.25	-	-
350	2.13	2.25	2.41	0.31
325	-	2.25	-	-
300	1.96	2.19	2.25	0.050
275	-	2.11	-	-
258	-	2.01	2.08	-
250	1.78	-	1.40	0.011
225	-	1.90	1.00	0.0073
215	-	1.78	-	-
210	-	1.72	-	-
200	1.56	1.62	0.78	0.0042
175	-	1.44	0.62	.0025
150	1.33	0.93	0.51	.0015
130.3	1.24	0.183	0.45	-
130.3	0.0004	0.0010	0.0006	.00104
129.3	-	-	-	.000026
120	0.00025	.0008	0.0004	-

t	κ (mhos)	t	κ (mhos)
	0%		
650	2.47	350	2.14
600	.43	300	1.97
554	.36	250	1.75
550	.64	200	1.57
500	.52	150	1.33
450	.41	144.6	1.31 - 0.00034
400	.28	140	0.00026

Silver chloride ($AgCl$) + Silver sulfide (Ag_2S)

Truthe, 1912

%	f.t.	E	min.	tr.t.	min.
0	455	-	-	-	-
8	446	376	2.3	-	-
15	435	376	3.1	-	-
22	422	373	4.7	177	6.0
30	412	376	7.6	175	6.3
40	397	375	15.8	-	-
50	-	376	20.5	175	7.6
65	571	375	16.0	176	8.0
80	767	374	5.0	175	8.2
92	821	376	1.5	176	-
100	834	-	-	175	12.1

Sandonnini, 1912

mol%	f.t.	E	min.
0	455	-	-
5	440	-	-
10	422	-	-
20	400	380	30
30	390	380	50
35	E	380	180
40	400	378	100
50	480	378	80
60	596	375	50
70	630	380	40
80	698	372	20
90	750	-	-
95	782	-	-
100	836	-	-

Urazov and Chelidze, 1941

wt%	f.t.	E	min.
0	455	-	-
15	431	363	3
20	416	362	4
30	394	363	6
35	382	365	8
40	378	365	10
50	-	365	14
60	514	365	12
65	549	364	9
70	581	362	7
75	656	362	6
100	834	-	-

Silver chloride (AgCl) + Silver nitrate ($AgNO_3$)
Kablukov, 1908

mol%	f.t.	mol%	f.t.
100	207.2	90.47	190.6
98.53	204.9	89.13	187.8
96.75	201.8	87.54	185.0
93.84	197.3	85.01	179.8
92.67	195.0		

Scarpa, 1912

mol%	f.t.	E	min.	tr.t.
0.00	454	-	-	-
5.00	444	155	50	80
20.00	395	155	70	82
30.00	360	158	115	85
40.00	324	160	180	"
50.00	278	162	210	"
60.00	232	160	260	"
70.00	185	"	280	"
75.00	170	"	300	"
80.00	172	"	260	84
85.00	185	"	150	83
90.00	195	"	30	85
95.00	205-175	-	-	105
100.00	210	-	-	155

Lifshits, 1956.

%	f.t.	%	f.t.
100	208	69	200(1+1)
81.5	176 E	0	452

Spooner and Wetmore, 1951

mol%	d 310°	d 320°	d 330°
100	0.3857	0.3847	0.3837
89.51	.3932	.3922	.3912
85.52	.3969	.3959	.3949
82.13	.3987	.3977	.3969
77.13	.4028	.4018	.4008
66.30	.4121	.4111	.4101
55.85	.4218	.4209	.4200

t	κ(mhos) 100 mol%	89.51	85.52
210	(0.660)	-	0.663
220	.706	0.705	.710
230	.751	.752	.755
240	.796	.797	.800
250	.840	.841	.844
260	.883	.884	.888
270	.926	.926	.931
280	.968	.968	.973
290	1.010	1.009	1.015
300	.051	.050	.056
310	.092	.091	.097
320	.132	.132	.138
330	.173	.173	-

t	κ(mhos) 82.13 mol%	77.13	66.30	55.85
210	0.664	-	-	-
220	.712	-	-	-
230	.758	0.775	-	-
240	.803	.821	-	-
250	.847	.865	0.916	-
260	.890	.908	.961	-
270	.932	.950	1.005	1.092
280	.974	.992	.049	.136
290	1.015	1.035	.093	.180
300	.056	.077	.137	.224
310	.097	.119	.181	.268
320	.138	.161	.223	.312
330	-	.203	.266	.356
340	-	.244	.308	-
350	-	-	.349	-

Silver chloride (AgCl) + Silver tungstate (Ag_2WO_4)

Lesnykh and Bergman, 1956.

mol%	f.t.	mol%	f.t.
100	584	17.7	390°E
33.4	455	0	455

Silver bromide (AgBr) + Silver iodide (AgI)

Mönkemeyer, 1906

%	f.t.	tr.t.	%	f.t.	tr.t.
0	422	-	65.21	434	-
12.19	397	-	74.52	462	96
23.50	381	-	83.33	490	103
34.87	378	-	91.83	520	114
45.57	389	-	100	552	143
55.55	412	-			

Tubandt and Lorenz, 1914

mol%	f.t.	m.t.	tr.t.	E
0	420	-	-	-
10	400	-	-	-
20	380	-	-	128
30	375	375	340	"
40	390	380	300	"
50	410	395	265	"
60	430	412	230	"
70	465	438	190	"
80	490	470	150	"
90	525	500	125	"
92	-	-	130	-
100	552	-	143	-

Stasin and Teltov, 1949 (fig.)

mol%	f.t.	m.t.	tr.t. I	tr.t. II	E
0	430	430	-	-	-
10	405	400	-	-	-
18	385	370	-	370	-
20	380	372	-	350	-
21	375	375	-	-	-
30	380	380	350	240	-
40	395	390	310	180	-
50	415	405	270	127	127
60	440	430	230	-	"
70	465	460	190	-	"
80	495	485	150	-	"
85	-	-	127	-	"
90	-	-	127	-	132
100	560	560	140	-	140

Tubandt and Lorenz, 1914

t	ϰ (mhos) 95 mol%	90	80	70	60
600	2.46	2.42	2.45	2.47	-
550	.40	.39	.41	.43	2.50
540	.39	-	-	-	-
529	.50	-	-	-	-
519	-	2.37	-	-	-
500	2.42	-	2.36	2.39	2.44
495	-	2.47	-	-	-
490	-	-	2.38	-	-
465	-	-	-	2.35	-
463	-	-	2.43	-	-
450	2.29	2.34	2.39	-	-
442	-	-	-	2.43	2.35
434	-	-	-	-	-
414	-	-	-	-	2.33
402	-	-	-	-	2.45
400	2.16	2.20	2.25	2.33	-
300	1.87	1.88	1.94	2.00	2.04
250	1.70	1.70	1.76	1.82	1.84
229	-	-	-	-	1.75
225	-	-	-	-	1.15
200	1.50	1.50	1.56	1.62	0.67
196	-	-	-	1.61	-
180	-	-	-	1.03	-
175	-	-	-	-	0.43
160	-	-	-	0.68	-
150	-	1.28	1.33	-	0.26
146	-	-	1.31	-	-
140	-	-	0.85	0.48	-
135.2	1.20	-	-	-	-
133	0.00024	-	-	-	-
130	0.00020	-	-	-	-
128	-	1.16	0.62	0.43	0.17
127.1	-	0.00084	.0015	.0029	.0018
120	0.00010	.00055	.0012	.0020	.0013
100	-	.00039	.0006	.0011	.00061
75	-	.00022	.00025	.00043	.00024

	40 mol %	30 mol %	20 mol %	10 mol %
600	2.70	2,72	2.74	2.90
550	.64	.67	.68	.84
500	.57	.60	.61	.77
450	.49	.52	.52	.67
402	-	-	-	.56
400	2.38	2.43	2.42	2.20
392	.37	-	-	-
385	.46	-	-	-
381	-	2.39	-	-
380	-	-	2.38	0.72
378	-	2.50	-	-
370	2.39	.49	-	-
360	-	.46	1.13	0.36
350	2.31	.42	-	-
346	-	.40	-	-
340	-	1.86	0.68	0.20
330	2.26	-	-	-
320	-	1.11	0.40	0.11
305	2.17	-	-	-
300	1.60	0.68	0.22	0.060
275	0.90	0.38	.10	.022
250	0.55	.22	.048	.011
225	.34	.13	.025	.0045
200	.22	.083	.010	.011
175	-	-	.0046	.00014
150	0.11	0.040	.0020	.00022
130	-	-	.00096	.000095
128	0.091	0.031	0.00080	-
127.1	.0030	.00048	.00024	-
120	.0012	.00038	.00017	-
100	.00030	.00025	.00006	-
75	.00006	-	-	-

Tammann, 1915 sat.t. annealed

%	for 40-60 hours at 300°	for 13 days at 280°
100	141.4	-
91.32	115	-
83.30	113	-
74.49	112	107
64.79	113	106
60.34	-	103
56.04	111	104
50.60	-	103
45.44	112	101
34.18	112	-
23.95	110	-
20.93	112	-

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ 500°	σ 600°	mol%	σ 500°	σ 600°
0	153	150	75	122	119.5
25	139.5	136	100	117.5	115.5
50	130	125.5			

Barth and Lunde, 1926

Lattice constants (see authors)

Silver bromide (AgBr) + Silver nitrate ($AgNO_3$)

Kablukov,1908

mol%	f.t.	mol%	f.t.
100	207.2	79.80	163.5
98.99	205.6	74.26	166.0
95.04	199.2	70.55	175.0
90.52	190.4	67.92	179.0
88.88	186.6	65.96	182.8
87.62	183.8	63.91	186.0
84.94	177.0	59.25	191.2
81.85	169.0	54.96	194.0
81.25	167.0	39.39	222.0

Scarpa, 1913

mol%	f.t.	E	min.	tr.t.
0.00	420	-	-	-
5.00	400	180	50	-
20.00	342	185	100	-
30.00	302	185	150	-
40.00	265	188	210	-
50.00	210	188	240	-
55.00	190	145	20	80
60.00	185	154	60	"
70.00	175	155	230	"
75.80	-	"	280	"
80.00	165	"	200	82
90.00	195	148	60	85
95.00	205	-	-	90
97.50	208-190	-	-	110
100.00	210	- (1+1)	-	155

Silver iodide (AgI) + Silver nitrate ($AgNO_3$)

Hogervorst, 1906

mol%	f.t.	tr.t.	mol%	f.t.	tr.t.
100	209.5	159.1	62.5	117.9	-
97.6	205.6	159.2	58.0	113.2	-
96.3	202.5	159.2	-	109 E	-
93.3	195.9	159.4	50	111.5(1+1)	-
89.9	185.8	160.2	43.64	109 "	-
86.1	171.5	160.2	41.5	106.3 E	-
84.0	-	160.4	-	105.5	-
82.8	155.4	160	39.9	112.4	-
80.0	144.5	-	38.75	121.6	116-134
77.4	131.8	-	38.0	127.6	" "
75.0	117.8	-	33.4	212.8	" "
74.5	114.7 E	-	25.0	331	-
72.6	116.2	-	15.0	420	-
66.7	119.1	(1+2)	0.0	526	146

Kablukov, 1908

mol%	f.t.	mol%	f.t.
100	207.2	95.47	198.4
99.19	205.8	93.46	194.0
97.59	203.4	88.36	180.4

Scarpa, 1913

mol%	f.t.	E	min.	tr.t.	min.
0.00	548	-	-	140	300
2.50	540	-	-	125	280
5.00	525	-	-	105	250
10.00	470	65	120	105	225
20.00	375	75	350	105	150
30.00	265	78	400	100	80
40.00	140	80	430	105	25
45.00	E	80	480	-	-
50.00	100	80	100	-	-
55.00	105	80		-	-
60.00	108	-	-	-	-
65.00	-	-	-	-	-
70.00	115	105	350	-	-
75.00	130	105	300	-	-
80.00	148	105	280	-	-
85.00	170	105	250	-	-
90.00	185	102	200	-	-
95.00	200-165	92	100	-	-
97.50	210-190	-	-	-	-
100	210	-	- (1+2)	-	-

Dombrovskaya and Koloskova, 1953 (fig.)

Dombrovskaya, 1953 (fig.)

%	f.t.	%	f.t.
0	554	67	119 (1+2)
40	100	73	110
50	109 (1+1)	100	208
58	106		

Lifshits, 1956.

%	f.t.	%	f.t.
100	208	54.5	188(1+1)
76	155 E	0	419

Bokhovkin, 1950

%	d 225°	200°	175°	150°
100	3.922	-	-	-
90	4.108	4.164	-	-
80	.225	4.346	4.398	4.418
75	.368	4.442	.497	.543
70	.466	4.538	.605	.638
65	.563	4.629	.698	.727
60	.652	4.720	.782	.826
55	.742	4.815	.873	.921
50	.835	4.902	.965	5.013
45	.929	4.993	5.054	5.102
40	5.023	5.100	5.141	5.193

%	225°	200°	η 175°	150°	125°
100	4560	-	-	-	-
90	5120	5620	-	-	-
80	5640	6430	7550	9410	-
75	5960	7510	8610	10700	15700
73.5	6030	7600	8850	11000	16090
70	5830	7420	8630	11460	17000
67.5	6210	7830	9270	11010	17730
65	6650	8140	9420	11230	17090
60	7320	8730	10500	13780	19760
55	8210	9950	11790	15050	21920
52	8090	9680	11990	15730	22740
50	7980	9860	11720	16080	23200
48	8610	10690	12180	14990	29430
45	9150	11140	12890	16090	22560
40	9680	11860	14460	19360	30570

%	300°	275°	250°	ϰ 225°	200°	175°	150°	125°
100	9930	9070	8040	6910	-	-	-	-
90	9800	8820	7840	6690	5700	-	-	-
80	9930	9000	8000	7050	5930	5000	3880	-
75	10100	9230	8310	7350	6300	5280	4180	3130
73	10000	9150	8240	7260	6200	5220	4160	3200
70	10200	9270	8310	7380	6300	5280	4260	3220
67	10300	9430	8560	7650	6560	5540	4510	3480
65	10400	9550	8670	7780	6720	5700	4700	3680
63	10500	9760	8850	7910	6830	5840	4780	3820
60	10700	9340	9070	8070	7000	6080	5090	4000
57	11000	10200	9350	8350	7350	6350	5240	4240
54	11300	10400	9630	8740	7650	6540	5610	4580
52	11400	10300	9590	8700	7560	6590	5580	4610
50	11500	10700	9930	9040	7940	7110	5980	4910
47	11900	10900	10100	9230	8350	7260	6280	5010

Silver sulfide (Ag_2S) + Silver selenide (Ag_2Se)

Panebianeo, 1915

%	f.t.	m.t.	tr.t.
0	839	-	177
2.9	820	806	-
8.0	800	780	155
11.9	798	773	-
24.2	773	769	-
27.0	771	767	115.5
28.3	774	770	112
30.5	775	771	103
38.9	792	773	-
47.2	808	775	93.5
49.9	815	798	-
54.2	825	783	84.0
60.5	822	781	80.0
67	830	821	-
74.2	842	795	74.5
83.1	-	-	82
85	848	827	-
92.7	856	840	88.5
98.5	854	850	-
100	855	-	111.5

%	d	%	d
0	7.253	60.5	7.77
8	.29	81	7.97
27	.44	83.1	8.05
47.2	.655	100	8.138
54.3	.755		

Proustite (Ag_3AsS_3) + Pyrargyrite (Ag_3SbS_3)

Jaeger and van Klooster, 1912

%	f.t.	%	f.t.
0	489.5	52.2	475
5.4	487.5	62.1	473.5
10.8	485	71.8	475
16	483	81.4	475.5
21.4	476	90.8	481.7
31.9	476	95.1	402.2
42.1	476	100	483

Silver sulfate (Ag_2SO_4) + Silver molybdate (Ag_2MoO_4)

Belyaev and Doroshenko, 1954

mol%	f.t.	mol%	f.t.
100	554	50	526
95	550	45	535
90	541	40	544
85	534	35	555
80	524	30	564
75	514	25	576
70	505	20	590
67.5	498	15	603
65	500	10	620
62.5	505	5	637
60	509	0	658
55	519		

Belyaev and Doroshenko, 1956.

mol%	f.t.	mol%	f.t.
100	555	55	510
95	546	50	518
90	537	45	528
85	528	40	538
80	519	35	552
75	510	30	556
72.5	506	25	580
70	503	20	595
67.5	500 E	15	610
65	501	10	624
62.5	502	5	639
60	505	0	654

Magnesium fluoride (MgF_2)
+ Magnesium orthophosphate ($Mg_3O_8P_2$)

Winter, 1913

%	f.t.	E	min.	tr.t.
0	1221	-	-	-
10	1204	1070	25	790
15	1192	1083	35	833
20	1176	1105	45	853
30	1140	1103	110	871
40	1120	1101	160	868
50	1179	1068	65	880
60	1223	1002	33	884
70	1247	-	-	810
76.3	1253	-	-	838
80	1243	-	-	952-783
90	1172	1076	30	972
92.1	1143	1076	15	979
96.7	1120	1076	25	975
97.3	1152	1058	15	963
100	1184	-	-	-

(1+1)

Magnesium chloride ($MgCl_2$) + Magnesium sulfate ($MgSO_4$)

Jänecke, 1912

mol%	f.t.	E	mol%	f.t.	E
100	1185	-	30	808	667
60	1040	650	20	-	"
50	950	673	10	708	"
40	908	664	0	703	-

Speranskaya, 1938

mol%	f.t.	mol%	f.t.
0	712	20.9	661
5.3	696	22.4	657
10.5	685	23.8	671
13.8	676	25.1	696
16.6	672	26.5	712
19.1	664		

E : 22.5 mol% 656°

Magnesium oxide (MgO) + Magnesium sulfide (MgS)

Mikulinski and Kamkin, 1939 (fig.)

%	m.t.	f.t.	%	m.t.	f.t.
0	2500	2500	76	1590	1690
14	1980	2310	81	1590	1730
32	1590	2080	95	1810	1860
58	1610	1720	100	1900	-

Magnesium metasilicate ($MgSiO_3$)
+ Magnesium orthosilicate (Mg_2SiO_4)

Deleano, 1913

mol %	f.t.	mol %	f.t.	mol %	f.t.
80	1640	30	1480	10	1505
60	1605	20	1500	0	1500
40	1520				

Magnesium sulfate ($MgSO_4$). 7 H_2O
+ Magnesium chromate ($MgCrO_4$). 7 H_2O

Fock, 1880

mol%	n_D			
	α	β	γ	2V
0	1.4319	1.4549	1.4602	51.28
5.2	.4353	.4579	.4635	53.32
15.8	.4388	.4618	.4666	53.58
18.5	.4408	.4632	.4697	55.40
31.6	.4457	.4727	.4844	57.16
35.9	.4543	.4778	.4881	60.14
43.7	.4632	.4934	-	69.52
100	.5221	-	1.5680	75.28

Magnesium ammonium sulfate ($MgN_2H_8S_2O_8$)
+ Magnesium ammonium chromate ($MgN_2H_8Cr_2O_8$)

Porter, 1925

% L	% C	% L	% C	% L	% C
0	0	40	50	80	89
10	10	50	62	90	95
20	26	60	72	100	100
30	39	70	81		

vol%	n_D		
	α	β	
	crystal axes		
1.96	1.475	1.476	1.482
10.01	.491	.493	.495
26.21	.520	.525	.528
43.76	.554	.557	.560
62.79	.589	.590	.598
88.40	.627	.627	.641

Potassium magnesium sulfate $K_2SO_4.MgSO_4.6$ aq. +
Ammonium magnesium sulfate $(NH_4)_2SO_4.MgSO_4.6$ aq.

Retgers, 1890

%	d	%	d	%	d
0	2.039	40.90	1.893	73.36	1.797
13.48	1.984	51.59	1.860	90.43	1.755
20.29	1.962	58.52	1.827	100	1.721
28.75	1.935				

Magnesium cesium sulfate($MgCs_2S_2O_8.6H_2O$) +
Magnesium ammonium sulfate ($Mg(N_2H_8)_2S_2O_8.6H_2O$)
Wulff, 1902 (fig.)

%	extinction angle	d
100	6	1.72
76.8	14	.87
70.5	18	.92
60.4	22	2.00
51.2	30	.08
48.7	39	.20
29.5	47	.295
16.5	55	.445
9.4	60	.54
4.3	60	.60
0	62	.67

Ammonium magnesium chromate ($(N_2H_8)_2Mg(Cr_2O_8)_2$ +
Rubidium magnesium chromate $Rb_2Mg(CrO_4)_2$

Porter, 1925

% L	% C	% L	% C
0	0	60	60
20	20	80	80
40	40	100	100

%	d	%	d
100	2.463	53.05	2.122
96.88	.437	46.20	2.080
89.05	.374	31.97	1.998
81.05	.313	19.76	.932
71.06	.241	7.14	.869
61.05	.174	0	.835

vol%	n_D α	n_D β	n_D γ
100	1.6216	1.6330	1.6439
95.86	.6226	.6336	.6438
85.83	.6239	.6349	.6437
76.12	.6255	.6360	.6442
64.65	.6268	.6366	.6444
53.98	.6285	.6369	.6455
45.70	.6303	.6371	.6468
39.00	.6310	.6369	.6471
25.95	.6330	.6368	.6493
15.50	.6346	.6367	.6507
5.42	.6365	.6365	.6522
0	.6363	.6371	.6528

Calcium fluoride (CaF_2) + Calcium chloride ($CaCl_2$)

Plato, 1907

wt%	mol%	f.t.	E I	E II	min.
100	100	773.9	-	-	-
96.2	94.7	736.7	-	-	-
92.6	89.8	702.9	639.6	-	9
89.3	85.4	676.3	644.4	-	23
86.2	81.4	648.9	644.4	-	36
83.3	77.8	658.7	644.4	-	30
76.9	70.1	701.5	640.3	-	16
66.7	58.5	745.2	633.7	736.7	-
64.5	56.0	750.4	628.0	"	-
62.8	54.0	772.3	-	"	-
58.8	50.1	818.5	622.3	"	33
57.1	48.3	847.9	-	"	32
55.6	46.8	870.6	610.3	"	31
52.6	43.8	904.4	595.7	"	-
50.0	41.3	932.7	-	"	28
40.0	31.9	1025.8	-	"	22
30.0	23.2	1119.7	-	"	16

Bukhalov and Bergman, 1951 and 1952

mol%	f.t.	mol%	f.t.
100	773	68	713
96	742	64	725
92	720	60	730
88	696	56	788
80	658	53	820
76	687	52	833
72	700	48	1000
		0	1411

E : 81 mol% 650° tr.p. : 57.5% 735°

(1+1)

Calcium fluoride (CaF_2) + Calcium iodide (CaI_2)

Mc Creary, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	1414	82.5	668
47	1045	90	740
60	935	100	784
80	700		

Calcium fluoride (CaF_2) + Calcium oxide (CaO)

Budnikov and Tresvyatski, 1953

% initial	% final	f.t.	% initial	% final	f.t.
100	100	2570	30	40	1375
75	83.1	1700	20	30	1370
70	78.1	1550	18	27	1365
60	69	1430	15	23.5	1360
50	60	1390	10	17.3	1365
40	51	1380	0	0	1380

Bäak, 1954 (fig.)

mol%	f.t.	sat.t.	mol%	f.t.	sat.t.
0	1418	-	7.5	-	1485
0.8	-	1385	9.5	-	1385
5	-	1465	13.5	1380	-

Calcium fluoride (CaF_2) + Calcium orthophosphate ($Ca_3O_8P_2$)

Nacken, 1912

%	f.t.	E	min.
0	1392	-	-
10	1356	1200	80
19	1316	1203	150
27.2	1270	1205	250
43.4	1255	1205	250
60	1395	1198	170
76	1524	1180	100

Calcium fluoride (CaF_2) + Calcium silicate ($CaSiO_3$)

Karandeeff, 1910

%	f.t.	E	min.
100	1501	-	(70.6)
93.1	1381	1120	9.9
85.6	1293	1137	33.1
77.7	1196	1142	43.8
69.1	1128	1126	65.4
66.7	1127	1128	86.4
59.9	1160	1127	87.4
55	1210	1126	62.1
49.8	1302	1130	61.2
33.2	1338	1125	27.9
27.1	1363	1143	33.9
14.2	1363	1138	15.0
0	1378	-	(55.8)

Bäak and Ölander, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	1418	50	1240
4.0	1407	60	1310
9.1	1390	70	1380
20	1340	80	1440
30	1255	90	1495
41	1127 E	100	1540

Calcium chloride ($CaCl_2$) + Calcium bromide ($CaBr_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	760	60	740
90	745	50	741
80	742	40	743
70	741	0	780

Calcium chloride ($CaCl_2$) + Calcium iodide (CaI_2)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	740	50	555
90	720	40	570
80	685	38	575
70	650	0	780
60	610		

Calcium chloride ($CaCl_2$) + Calcium metasilicate ($CaSiO_3$)

Karandeeff, 1910

%	f.t.	E	min.	E_2
100	1501	-	(70.6)	-
90	1464	761	-	-
80	1390	-	23.4	-
70	1292	766	26.3	-
60	1213	759	38.0	-
50	1170	746	60.1	-
40	1127	754	68.7	705
30	1069	757	81.6	-
20	-	761	85.7	699
15	-	760	86.6	700
10	-	765	96.2	708
5	-	766	120.5	709
2	-	765	111.9	-
0	772	-	(125.0)	-

Calcium chloride ($CaCl_2$) + Calcium orthophosphate ($Ca_3P_2O_8$)

Nacken, 1912

%	f.t.	%	f.t.
3	786	12	1040
5	861	18	1160
7	900	30	1280
8	958		

Calcium chloride ($CaCl_2$) + Calcium sulfate ($CaSO_4$)

Golubeva and Bergman, 1954

mol%	f.t.
100	(1450)
12.5	708 E
0	774

Calcium oxide (CaO) + Calcium carbide (CaC_2)

Flusin and Aall, 1935

mol %	f.t.	mol %	f.t.
100	2310	50	1980
90	2200	40	1880
80	2040	36.6	1800 E
70	1800	30	2000
68	1750 E	20	2210
60	1920		

(1+1)

Aall, 1939 (fig.)

%	f.t.	%	f.t.
100	2320	40	1900 (1+1)
90	2195	35	1800 E
80	2025	30	2000
70	1805	20	2225
68	1750 E	10	2395
60	1930	0	2510
50	1975		

Calcium sulfide (CaS) + Calcium metasilicate ($CaSiO_3$)

Lebedev, 1911

mol%	f.t.	m.t.	sat.t. mixed crystal
100	1512	1512	-
95	1482	1472	1222
90	1482	-	1230
86.2	1496	-	-
85	1489	-	1301
80	1421	1404	-
75	1443	1404	1290
70	1443	-	1200
63	1432	1364	1148
50	1398	1370	-

Glaser, 1926

%	f.t.	E	%	f.t.	E
100	1512	-	90	1475	-
98	1500	-	80	1515	-
97	1480	1440	70	1520	-
95	1440	1440	60	1520	-
94	1470	1440	50	1545	-
92	1485	-			

Calcium tetraborate (CaB_4O_7) + Calcium sulfate ($CaSO_4$)

Jilenko and Sverkov, 1939

mol%	f.t.	mol%	f.t.
0	960	40	1103
5	954	50	1128
7	951	55	1115
10	948	60	1094
13	960	63	1082
15	980	65	1070
20	1003	70	1094
25	1035	80	1148 (1+1)
30	1063		

E_1 : 11 mol% 946 E_2 : 65 mol% 1070

Calcium carbonate ($CaCO_3$) + Calcium orthosilicate (Ca_2SiO_4)

Eitel, 1923 (fig.)

%	f.t.	E	tr.t.
0	1290	-	975
32	1180	1180	975
75.77	-	1180	1375 (1+2)
100	2130	-	1375

Calcium metasilicate ($CaSiO_3$) + Calcium metatitanate ($CaTiO_3$)

Smolenski, 1912

mol%	f.t.	tr.t.	mol%	f.t.	tr.t.
0	1512	-	43.8	1436	1329
11.5	1444	-	53.2	1452	1350
27.6	1430	1186	63.7	1472	1246
33.4	1420	1315	73.2	1484	-

Calcium metasilicate ($CaSiO_3$) + Calcium magnesium metasilicate ($CaMgO_6Si_2$)

Shairer and Bowen, 1942

%	f.t.	E	tr.t. I	tr.t. II	tr.t. III
0	1540	-	-	1125	-
2	1535	-	-	1140	-
5	1529	-	1360	1190	-
10	1517	-	1365	1260	1200
15	1500	-	"	1310	1290
17	1495	-	"	1335	1310
20	1490	-	"	1355	1340
23	1480	1355	"	1360	-
30	1460	"	"	-	-
40	1430	"	"	-	-
50	1400	"	"	-	-
55	1380	"	"	-	-
58	1370	"	"	-	-
60	1370	"	-	-	-
62	1360E	"	-	-	-
63	1360	"	-	-	-
65	1365	"	-	-	-
70	1370	"	-	-	-
80	1380	"	-	-	-

Calcium orthosilicate (Ca_2SiO_4) + Calcium phosphate ($Ca_4P_2O_9$)

Blome, 1910

%	f.t.	%	f.t.
100	1870	29.81	1780
80	1550	34.67	1780
60	1650	51.49	1680
40	1780	67.92	1710
20	1750		

Strontium fluoride (SrF_2) + Strontium chloride ($SrCl_2$)

Plato, 1907

wt%	mol%	f.t.	E	min.	
100	100	871.0	-	-	
98	97.5	842.1	723.8	-	
96.2	95.2	810.3	735.4	-	
92.6	90.8	770.4	751.4	18	
89.3	86.9	776.2	"	26	
86.2	83.2	813.9	"	23	
83.3	79.8	848.3	"	20	
80.7	76.8	880.3	"	14	
75.8	71.3	913.7	-	-	
70.0	64.9	932.7	717.1	-	
60.0	54.3	955.0	-	-	(1+1)
50.0	44.2	954.0	945.2	-	
45.0	39.3	948.4	"	-	
40.0	34.5	973.4	"	19	
35.0	29.9	1032.1	"	16	
30.0	25.3	1039.0	"	14	
20.0	16.5	1195.0	"	9	

Winter, 1913

%	f.t.	E	min.	
100	874	-	-	
95	816	753	85	
90	759	"	170	
80	879	"	110	
56	962	-	-	(1+1)
50	959	952	120	
40	983	947	210	
20	1193	934	110	
0	1400	-	-	

Bukhalov and Bergman, 1952

mol%	f.t.	mol%	f.t.
100	868	64	937
92	784	60	944
88	806	56	948
84	846	52	958
80	877	48	958
76	888	44	950
72	913	40	963
68	924	36	1000

E_1 : 86 mol% 763° E_2 : 42 mol% 944°

(1+1) 960°

Strontium fluoride (SrF_2) + Strontium orthophosphate ($Sr_3P_2O_8$)

Winter, 1913

%	f.t.	E	%	f.t.	E
0	1400	-	35	1281	1205
10	1356	1190	40	1324	1180
20	1275	1204	90.8	1685	- (1+1)
23	1237	"	100	-	-
30	1239	1207			

Strontium chloride ($SrCl_2$) + Strontium nitrate (SrN_2O_6)

Gromakov and Gromakova, 1953

mol%	f.t.	m.t.	mol%	f.t.	m.t.
100	645	-	50	474	433
80	560	465	40	525	426
70	495	438	30	560	427
60	442	424	0	872	872
56.5	428	(1+1)			

Tokareva and Bergman, 1956.

mol%	f.t.	mol%	f.t.
54	497	46	488
52	492	44	498
50	488	42	506
48	485	40	517
47	482 E	38	528

Strontium chloride ($SrCl_2$) + Strontium orthophosphate ($Sr_3P_2O_8$)

Winter, 1913

%	f.t.	E	min.	
0	874	-	-	
3	850	828	50	
10	893	"	95	
20	1030	819	85	
30	1139	812	75	
40	1239	806	55	
88.3	1625	-	-	(1+1)

Strontium chloride ($SrCl_2$) + Strontium sulfate ($SrSO_4$)

Golubeva and Bergman, 1955 (fig.)

%	f.t.
0	873
14	762 E
25	895

Strontium hydroxide (SrO_2H_2) + Strontium nitrate (SrN_2O_6)

Wolf, 1935

%	f.t.		%	f.t.
0	535		50	495
10	510		60	460
12	500		68	420
20	525	(4+1)	80	530
30	535		100	620
40	515			

Barium fluoride (BaF_2) + Barium chloride ($BaCl_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100.0	960	77.2	925
94.4	930	73.7	940
89.4	890	70.6	960
87.1	870	67.8	975
86.0	865	65.2	985
84.9	870	62.8	1000
83.9	880	0.0	1280
80.9	910		

Plato, 1907

wt%	mol%	f.t.	E	min.	
100	100	958.9	924.5	-	
99.5	99.4	951.1	"	-	
98.5	98.2	932.9	"	-	
98.0	97.6	924.8	"	-	
97.1	96.6	915.6	-	-	
96.2	95.5	907.6	-	-	
92.6	91.3	883.1	843.1	5	
89.3	87.5	864.1	846.3	18.5	
87.7	85.7	855.4	"	25	
86.2	84.0	848.7	"	32.5	
84.7	82.3	848.7	"	31	
83.3	80.7	857.0	"	27	(1+1)
80.7	77.9	885.7	"	21	
78.1	75.0	910.0	"	7	
70.0	66.2	975.4	843.1	-	
60.0	55.8	1001.5	834.0	-	
50.0	45.7	1005.0	927.0	-	
40.0	35.9	985.5	936.0	12	
30.0	26.5	945.2	942.0	28	
20.0	17.4	1058.1	942.0	22.5	
10.0	8.5	1193.7	942.0	12.5	

Winter, 1913

%	f.t.	E	min.
100	958	-	-
90	864	844	90
75	936	842	98
60	1001	831	20
54.3	1008	-	-
50	1003	933	35
40	981	"	110
20	1054	936	140
0	1289	-	-

(1+1)

Bukhalov and Bergman, 1951

mol%	f.t.	mol%	f.t.	
100	958	50	1008	(1+1)
90	900	40	997	
80	877	30	955	
76	913	26	936	
72	936	20	1017	
68	960	0	1280	
60	995			

E : 81% 854° 28% 936° tr.p. :71% 940°

Barium fluoride (BaF_2) + Barium bromide ($BaBr_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	880	83	850
95	840	80	915
90	800	0	1280

Barium fluoride (BaF_2) + Barium iodide (BaI_2)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	740	90	710
95 °	700	85	800
92	670	0	1280

Barium fluoride (BaF_2) + Barium orthophosphate ($Ba_3P_2O_8$)

Winter, 1913

%	f.t.	E	min.
0	1289	-	-
10	1258	1076	25
20	1185	1082	50
30	1106	1092	80
40	1167	1091	80
50	1280	1092	55
60	1373	1090	40
91.2	1670	-	-

(1+1)

Barium fluoride (BaF_2) + Barium metasilicate ($BaSiO_3$)

Bergman, Nesterova and Bichkova, 1955 (fig.)

mol%	f.t.	mol%	f.t.
0	1280	43	1075
35	1090	45	1100

(2+1) incongruent

Bychkova and Bergman, 1956.

mol%	f.t.		mol%	f.t.
32.5	1100		42	1075 E
35	1090	(2+1)	42.5	1078
37.5	1088		45	1100
40	1082			

Barium chloride ($BaCl_2$) + Barium bromide ($BaBr_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	880	76	890
90	870	0	960
80	885		

Barium chloride ($BaCl_2$) + Barium iodide (BaI_2)

Ruff and Plato, 1903

mol%	f.t.
100	740
90	700
80	725
0	960

Barium chloride ($BaCl_2$) + Barium orthophosphate ($Ba_3P_2O_8$)

Winter, 1913

%	f.t.	E	min.
0	958	-	-
10	905	894	145
15	954	898	165
20	1017	895	155
29.1	1100	895	140
39.1	1200	899	100
89.7	1584	-	-
(1+1)			

Barium chloride ($BaCl_2$) + Barium carbonate ($BaCO_3$)

Belyaev and Sholokhovich, 1952

mol %	f.t.	mol %	f.t.
0	964	11	890
1	950	12	886
2	932	13	884
3	922	14	877
4	917	15	872
5	913	16	867
6	910	17	862
7	906	18	870
8	898	20	880
9	897	23	903
10	894	26	920

E : 17.5 mol % 860°

tr.t. : 2 mol % 925°

Barium chloride ($BaCl_2$) + Barium metasilicate ($BaSiO_3$)

Voloskov , 1911

%	f.t.	E	%	f.t.	E
100	1464	-	40	1042	904
90	1371	902	30	1000	900
77	1225	904	20	972	899.5
65	1142	906	10	921	900
57	1104	900	5	936	903
50	1057	904	0	968	-

E : 2 mol% 902°

Barium chloride ($BaCl_2$) + Barium sulfate ($BaSO_4$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	955	30	900
10	915	35	900
19	900	40	925
20	900		

Barium oxide (BaO) + Barium carbonate ($BaCO_3$)

Lander, 1951

E : 67% 1030° tr.t. I 806° tr.t. II 968°

mol%	Q relat.* trans.I	trans.II	melting
96	140	24	0
90	133	21	15
82	118	18	50
69	107	17	95
67	102	16	100
61	91	16	98
50	68	12	80
40	48	8	59
31	41	5	44
20	24	3	26
6	2	0	6

* heat of transition or melting relative to that at the eutectic point (=100)

Barium sulfide (BaS) + Barium metasilicate ($BaSiO_3$)

Voloskov , 1911

mol%	f.t.	E	mol%	f.t.	E
100	1464	-	70	1374	1325
90	1404	1320	60	1474	1328
80	1368	1328	50	-	1325
75	1325	1325 E			

Barium metasilicate ($BaSiO_3$)
+ Barium metatitanate ($BaTiO_3$)

Smolenski, 1912

mol%	f.t.	mol%	f.t.
0	1470	42.1	1405
10.8	1461	52.2	1376
21.4	1450	62.1	1384
31.9	1420	71.8	1409

Barium metatitanate ($BaTiO_3$)
+ Barium metastannate ($BaSnO_3$)

Smolenski and Isupov, 1954

%	tr.t. I-II	tr.t. II-III	tr.t. III-IV
0	125	+22	-48
5	80	35	+5
10	45	40	+40
15	8	-	-
20	-25	-	-
25	-67	-	-
30	-110	-	-

Smolenski, 1956 (fig.)

mol%	tr.t. I	tr.t. II	tr.t. III
0	125	20	-50
3	98	27	-20
5	80	35	+ 2
8	60	40	22
10	40	40	40
15	15	-	-
20	-35	-	-
25	-70	-	-
30	-100	-	-

mol%	Curie point	mol%	Curie point
0	125°	15	12
5	85	20	-20
8	60	25	-73
10	50	30	-100

Barium metatitanate ($BaTiO_3$)
+ Barium metazirconate ($BaZrO_3$)

Smolenski and Isupov, 1954 (fig.)

mol%	tr.t.	mol%	tr.t.
0	125	25	-10
10	98	30	-50
15	70	35	-84
20	88	40	-120

Smolenski, 1956 (fig.)

mol%	Curie point	mol%	Curie point
0	125°	20	25
5	113	25	- 8
10	96	30	-50
15	60	35	-85

mol%	Lattice constant a	Lattice constant c
0	3.980	4.025
10	4.020	4.020
15	4.025	4.025
20	4.040	4.040
25	4.050	4.050
30	4.065	4.065
40	4.080	4.080
100	4.175	4.175

Lead fluoride (PbF_2) + Lead chloride ($PbCl_2$)

Sandonnini, 1911

mol%	f.t.	tr.t.	min.
100	495	-	-
97.5	484	-	-
95	477	453	30
92.5	465	454	60
90	E	454	70
80	525	454	60
70	570	452	45
60	591	451	25
55	597	425	10
50	601	-	-
45	537	-	-
40	592	553	8
30	566	554	40
25	E	554	80
22.5	568	556	60
20	570	554	20
15	630	570	150
10	694	567	120
5	758	566	50
0	824	-	-
(1+1)	(4+1)		

Pélabon and Laude, 1929

mol%	f.t.
100	498
91.4	461.5 E
0	908

Lead fluoride (PbF_2) + Lead oxide (PbO)

Sandonnini, 1914

mol%	f.t.	E	min.
0	824	--	-
5.00	800	483	-
11.00	755	485	60
21.53	701	490	90
31.57	632	492	130
42.21	565	494	160
45.00	545	494	220
52.34	-	"	250
56.00	-	"	280
62.20	658	"	230
65.00	588	493	210
72.02	658	"	180
81.44	718	492	130
91.35	848	490	80
95.00	848	490	40
100	892	-	-

Lead fluoride (PbF_2) + Lead phosphate ($Pb_3P_2O_8$)

Amadori, 1912

%	f.t.	E	min.
0	820	-	-
2.5	774	698	50
5	746	"	90
7.5	-	"	150
10	742	"	120
15	800	"	100
20	846	"	90
30	922	"	70
35	946	"	60
50	1032	695	40
60	1072	694	30
65	1084	690	20
75	1092	686	10
78	1098	-	-
80	1094	1000	30
85	1080	1004	70
90	1058	"	100
95	1032	"	140
100	1014	-	-

(1+1)

Lead fluoride (PbF_2) + Lead sulfate ($PbSO_4$)

Gladushchenko and Bergman, 1955

mol%	f.t.
0	860
26.5	520 E
100	1100

Gladushchenko and Bergman, 1956

mol %	f.t.	mol %	f.t.
10	728	35	637
15	668	40	687
18	630	43	718
20	605	46	746
23	571	49	775
25	542	52	800
26.5	520 E	55	824
30	571		

Lead chloride ($PbCl_2$) + Lead bromide ($PbBr_2$)

Jellinek and Golubowski, 1930

t	p_1	p_2
	25 mol%	
663	15.18	6.20
694	26.27	10.33
741	51.50	21.67
779	91.91	40.43
	50 mol%	
662	11.42	13.55
694	18.72	23.27
744	38.45	49.55
778	66.67	87.25
	75 mol%	
663	5.89	21.54
692	8.64	35.86
742	18.86	74.20
781	29.60	132.40

Greiner and Jellinek, 1933

mol% L	V	p_1	p_2
	770°		
0	0	101	0
10	9.5	84.3	8.77
30	21.5	72.1	31.0
50	52.8	59.8	66.7
70	69.1	40.1	91.6
90	86.6	17.5	122.7
100	100.0	0	145

Mönkemeyer, 1906

%	f.t.	%	f.t.
0	495	66.43	423
12.78	482	75.50	407
24.12	469	84.08	397
37.14	457	92.24	382
46.82	443	100	370
56.90	433		

Favorski , 1941 (fig.)

mol%	f.t.
0	498
30	460
70	410
100	373

Calingaert, Lamb and Meyer, 1949

mol %	f.t.	m.t.
0	496.3	496.0
12.5	478.9	474.6
25	464.0	457.5
35	452.3	443.3
50	435.9	425.0
62.5	420.7	410.2
75	403.4	394.2
87.5	386.5	381.5
100	370.1	369.8

Boardman, Dorman and Heyman, 1949

mol %	molar volume (in cc)
0	57.92
19.7	59.99
50.3	63.32
84.6	67.4
100	68.60

Sandonnini, 1913 and 1914

%	κ(in mhos)	%	κ(in mhos)
	500°		
100	1.030	30.56	1.310
92.24	.059	12.79	.400
79.85	.108	0	.472
56.91	.201		

mol%	κ 200°	250°	300°
100	0.50	1.40	6.0
85	.34	0.90	4.0
70	.22	.75	2.5
60	.16	.54	2.3
50	.12	.40	2.1
40	.17	.67	2.4
30	.38	1.02	3.4
15	.46	2.23	5.8
0	.81	2.74	8.6

Mulcahy and Heymann, 1943 (fig.)

mol%	λ
500°	
100	35
90	35.5
75	36
50	36
25	39
10	39.5
0	41

Harrap and Heymann, 1955

mol%	η	κ	η	κ
	425°		450°	
100	5680	0.782	4830	0.836
79.7	-	.818	-	.912
75.3	6260	-	5240	-
60.2	-	0.860	-	0.968
56.4	6800	-	5530	-
38.5	-	-	-	1.022
	475°		500°	
100	4220	0.952	3730	1.032
79.7	-	1.008	-	1.098
75.3	4500	-	3940	-
60.2	-	1.070	-	1.172
56.4	4680	-	4030	-
38.5	-	1.135	-	1.246
23.1	5200	-	4450	-
19.0	-	1.212	-	1.336
0	-	-	4560	.430
	525°		550°	
100	3320	1.112	2970	1.190
79.7	-	.182	-	.258
75.3	3470	-	3050	-
60.2	-	1.268	-	1.360
56.4	3520	-	3140	-
38.5	-	1.354	-	1.450
23.1	3850	-	3380	-
19.0	-	1.446	-	1.556
0	4010	.566	3540	.694

Salstrom and Hildebrand, 1930

t	e
100%	
438.3	1.0692
443.5	.0665
451.0	.0615
453.0	.0601
465.6	.0525
468.0	.0516
484.3	.0416
496.2	.0330
501.1	.0306
517.1	.0215
520.0	.0200
527.6	.0156
536.0	.0100
556.0	0.9979
561.2	.9945
576.0	.9856
80%	
450.7	1.0731
486.3	.0523
515.6	.0350
551.6	.0148
576.5	.0002
60%	
445.3	1.0888
468.5	.0765
493.9	.0626
514.2	.0514
531.4	.0422
535.5	.0400
552.6	.0304
578.9	.0160
583.6	.0136
50%	
450.1	1.0957
464.3	.0879
488.2	.0750
516.2	.0599
536.3	.0490
580.2	.0249
45%	
448.0	1.1053
475.7	.0897
501.3	.0756
507.5	.0730
528.4	.0616
553.9	.0468
584.9	.0308

Lead chloride ($PbCl_2$) + Lead iodide (PbI_2)

Greiner and Jellinek, 1933

mol% L	V	p_1	p_2
	760°		
0	0	88.6	0
10	17.3	95	20.0
30	40.1	85.6	57.7
50	60.2	68.5	102.2
70	77.2	45.2	152.5
90	93.9	12.9	197.9
100	100	0	228

Mönkemeyer, 1906

%	f.t.	E	min.
0	495	-	-
15.56	461	306	20
29.30	424	"	40
41.54	399	"	60
52.50	372	"	80
62.38	353	"	100
71.32	338	"	120
79.46	316	"	160
86.90	314	"	160
93.54	337	"	90
100	358	-	-

Pélabon and Laude, 1929 (fig.)

mol%	f.t.
0	498
38	338 E
100	380

Ilyasov and Bergman, 1956, and

Ilyasov and Bostandzhiyan, 1956.

mol%	f.t.	mol%	f.t.
0.0	496	50.0	362(1+1)
5.0	473	55.0	361
10.0	454	60.0	360
20.0	422	65.0	355
30.0	396	67.5	349
35.0	380	70.0	344
40.0	365	72.5	336
42.0	360 E	73.5	334 E
45.0	360	75.0	340
		80.0	358

Sumarokova and Modestova, 1956.

mol%	f.t.	tr.t.	L_1+L_2	E_1	E_2
0	501	422	-	-	-
5	479	-	-	328	-
5	460	-	-	334	-
10	454	-	-	335	-
10	454	-	355	329	-
15	435	-	356	324	-
20	421	-	362	331	-
20	421	-	351	331	-
25	398	-	363	329	-
25	402	-	360	-	-
30	388	-	359	329	-
33	375	-	365	329	-
40	363	-	363	329	-
40	-	-	364	-	-
45	-	-	361	-	-
50	-	-	358	-	308
50	-	-	360	322	-
50	-	-	358	-	306
55	-	-	356	331	-
60	353	-	-	328	308
60	350	-	-	324	308
65	345	-	-	329	-
70	332	-	-	-	306
70	-	-	-	-	-
72	-	-	-	328	-
75	332	-	-	330	308
75	332	-	-	-	308
75	331	-	-	322	308
77	331	-	-	329	306
80	348	-	-	-	308
80	339	-	-	330	308
85	358	-	-	328	309
90	371	-	-	329	307
90	379	370	-	-	308
95	388	-	-	322	308
95	391	-	-	322	312
98	390	-	-	-	308
98	387	365	-	-	306
100	403	372	-	322	-

Lead chloride ($PbCl_2$) + Lead oxide (PbO)

Baroni, 1934

%	f.t.	E	%	f.t.	E
0	(500)	-	62.0	695	-
10.0	472	438	65.0	688	-
20.0	438	"	67.5	698	-
25.0	478	439	70.0	708	-
32.5	515	438	73.0	715	-
37.5	535	"	76.0	723	-
41.0	570	439	78.5	722	710
45.0	612	525	82.0	710	"
50.0	645	526	87.0	760	"
55.0	673	525	94.0	805	"
60.0	690	"	100	(840)	-

(1+1) and (1+4)

Pélabon and Laude, 1929

mol%	f.t.
0	498
30	402 E
100	765

Ruer, 1906

%	f.t.	E I	min.	II	min.	tr.t.	min.
0	499	428	70	-	-	-	-
5	487	430	150	-	-	-	-
10	469	430	350	-	-	-	-
15	449	438	700	-	-	-	-
20	-	438	350	-	-	-	-
25	472	436	210	-	-	-	-
30	502	436	100	-	-	-	-
35	522	436	60	524	340	-	-
38	540	425	20	524	340	-	-
40	567	-	-	522	350	-	-
45	618	-	-	518	250	-	-
50	654	-	-	511	150	-	-
55	673	-	-	509	90	-	-
58	689	-	-	490	50	-	-
60	692	-	-	-	-	-	-
62	693	-	-	-	-	-	-
64	691	-	-	-	-	-	-
66	696	-	-	-	-	-	-
68	699	-	-	-	-	-	-
70	703	-	-	-	-	570	60
72	708	-	-	-	-	570	80
74	710	-	-	-	-	570	70
76	711	-	-	-	-	530	30
78	710	-	-	-	-	-	-
80	709	-	-	-	-	-	-
82	-	-	-	703	460	-	-
84	718	-	-	703	330	-	-
86	736	-	-	698	260	-	-
88	763	-	-	695	200	-	-
100	835	-	-	-	-	-	-

Lead chloride ($PbCl_2$) + Lead sulfide (PbS)

Truthe, 1912

%	f.t.	E	min.
0	499	-	-
5	483	441	-
10	470	440	1.3
15	463	441	5.8
20	442	442	16.1
30	554	439	9.5
45	677	446	6.5
60	831	442	5.0
75	953	448	1.9
88	1002	448	1.4
95	1044	450	0.6
100	1106	-	-

Lead chloride ($PbCl_2$) + Lead phosphate ($Pb_3P_2O_8$)

Amadori, 1912

wt%	mol%	f.t.	E	min.	
0	0	494	-	-	
12	4.44	-	480	220	
20	7.84	-	"	190	
30	12.81	792	"	170	
40	18.48	890	"	140	
50	25.55	974	"	120	
65	38.90	1062	478	110	
70	44.44	1090	475	90	
74.48	50.00	1110	474	80	
80	57.63	1130	"	50	
85	65.84	1145	470	20	
87.5	70.42	1150	468	10	
89.74	75.00	1156	-	-	(1+3)
92	79.77	1150	986	30	
94	84.31	1135	990	50	
95	86.68	1126	994	90	
97.5	93.05	1074	"	130	
98.5	95.75	1040	996	160	
100	100	1014	-	-	

Lead chloride ($PbCl_2$) + Lead vanadate (PbV_2O_6)

Bukhalov and Aleshkina, 1953 (fig.)

mol%	f.t.	mol%	f.t.
100	952	25	900
93.8	940 E	77	650
57	1005	21	476 E
44	990	0	498

(1+3) 1005°

Lead bromide ($PbBr_2$) + Lead iodide (PbI_2)

Mönkemeyer, 1906

%	f.t.	E	min.
0	370	-	-
12.25	348	251	60
23.91	325	"	100
34.99	302	"	125
45.57	278	"	165
55.67	262	256	225
65.32	278	"	205
74.55	296	254	195
83.40	319	252	160
91.87	339	251	100
100	358	-	-

Lead bromide ($PbBr_2$) + Lead oxide (PbO)

Sandonnini, 1914

mol%	f.t.	E	min.	addition compound t	addition compound min.
0	368	-	-	-	-
4.19	359	348	90	-	-
8.91	-	351	180	-	-
15.45	381	349	190	-	-
22.63	429	"	160	-	-
29.92	478	"	100	-	-
34.96	473	348	60	-	-
41.65	530	347	20	476	120
50.06	630	-	-	475	160
52.80	655	-	-	482	110
63.80	689	-	-	470	50
66.81	712	-	-	-	-
71.80	707	700	50	-	-
75.72	-	"	130	-	-
78.04	721	"	80	-	-
79.98	742	"	-	-	-
83.91	-	"	-	742	-
87.02	782	"	-	746	-
91.05	799	-	-	740	-
95.32	842	-	-	-	-
100	892	-	-	-	-

Baroni, 1934

%	f.t.	E	%	f.t.	E
0	(360)	-	55.0	722	-
2.5	355	340	59.0	720	700
7.0	345	339	64.0	712	699
12.5	375	340	68.0	700	700
19.0	430	"	71.5	710	"
25.0	465	"	75.0	720	"
27.5	475	"	78.5	725	-
30.0	527	339	84.0	750	725
33.0	575	"	91.0	800	"
40.0	648	475	96.0	842	"
50.0	710	"	100	(890)	-
(1+1)			(1+2)	(1+6)	

Knowles, 1951

mol%	f.t.	tr.t.	E
0	370	-	-
7.00	358	-	347
12.00	347	-	"
15.50	380	-	"
20.00	414	-	348
25.00	447	-	"
27.50	461	-	"
30.00	475	-	347
32.25	488	-	"
35.00	495	-	"
37.50	508	497	"
40.00	537	"	"
42.25	567	"	"
45.00	592	"	"
50.00	640	"	"
55.00	676	"	"
60.00	695	"	"
62.25	703	"	-
65.00	707	"	-
66.67	709	-	-
68.25	707	-	697
70.00	704	-	"
71.36	700	-	"
72.50	697	-	"
73.67	703	-	"
75.00	707	-	-
77.50	706	-	706
80.00	713	-	704
82.50	727	-	"
85.00	737	-	"
87.50	741	-	-
88.20	740	-	737
90.00	761	-	"
92.50	780	-	"
95.00	831	-	"
100.00	888	-	-

(1+2)

Lead iodide (PbI_2) + Lead oxide (PbO)

Baroni, 1934

%	f.t.	E
0	(400)	-
5.0	395	370
10.0	385	369
15.0	380	369
20.0	370	369
25.0	437	369
29.5	500	370
36.0	570	370
42.5	625	371
49.0	660	370
51.5	655	640
54.0	640	640
57.5	669	639
61.5	680	640
66.0	690	-
68.0	712	690
70.5	732	690
74.0	750	-
80.0	790	750
86.5	830	749
89.0	845	749
94.5	868	750
100	(890)	-
(1+2)	(1+4)	(1+6)

van Klooster and Owens, 1935

mol %	f.t.	m.t.	tr.t. I	tr.t. II	E
0	412	-	-	-	-
5.00	400	-	356	-	-
10.03	393	-	360	-	-
15.02	388	-	360	-	-
20.00	377	-	366	-	-
25.08	370	-	-	-	-
30.00	364	-	-	-	-
35.00	408	-	368	-	-
40.00	435	-	363	-	-
42.05	463	-	370	-	-
45.00	505	-	350	-	-
50.00	548	-	348	-	-
52.00	593	-	-	463	-
55.00	605	-	-	465	-
60.00	614	610	-	463	354
61.87	616	553	-	463	348
65.00	619	615	447	468	341
67.04	620	613	468	463	-
70.00	630	623	-	-	-
74.98	635	613	-	-	-
79.03	643	635	-	-	-
80.04	657	618	-	-	-
84.42	675	-	-	651	-
87.01	735	-	-	-	-
90.00	770	-	-	-	-

(2+3)

Lead selenide (PbSe) + Lead telluride (PbTe)

Elagina and Abrikosov, 1956, (fig.)

mol%	f.t.	m.t.
100	920	920
90	920	900
80	900	900
70	940	910
60	950	910
50	955	920
30	970	950
20	1010	970
10	1020	1010
0	1088	1088

mol%	lattice constant a (fig)	mol%	lattice constant a (fig)
100	6.4	40	6.18
90	6.38	30	6.13
80	6.3	20	6.11
70	6.29	10	6.10
60	6.25	0	6.09

Lead metatitanate ($PbTiO_3$) + Lead molybdate ($PbMoO_4$)

Belyaev, Sholokhovich, Barkova, 1956.

mol%	f.t.
100	1055
96.5	1019
72.5	1016

Lead metatitanate ($PbTiO_3$) + Lead metastannate ($PbSnO_3$)

Smolenskii, 1956.

mol%	tr.t.	mol%	tr.t.
70	200	40	350
60	250	30	400
50	300		

Lead metazirconate ($PbZrO_3$) + Lead metastannate ($PbSnO_3$)

Smolenskii, 1956. (fig.)

mol%	tr.t. I	tr.t. II
40	270	140
30	275	170
20	250	180
10	245	210
0	230	230

Lead sulfate ($PbSO_4$) + Lead chromate ($PbCrO_4$)

Jaeger and Germs, 1921

mol%	f.t.	tr.t. I	tr.t. II
100	844	783	707
90	850	816	748
80	861	813	"
70	879	844	-
60	906	856	-
50	937	868	-
45	959	847	-
40	985	874	934
35	"	"	"
30	"	"	"
25	1025	875	"
20	1073	873	930
10	-	"	911
3	-	-	877
0	1170	-	864

Lead sulfate ($PbSO_4$) + Lead molybdate ($PbMoO_4$)

Jaeger and Germs, 1921

%	f.t.	E	tr.t.
100	1065	-	1065
95	1059	1059	984
90	1051	950	979
80	1027	958	879
70	1000	962	"
60	971	963	"
55	968	962	"
50	985	"	878
40	1016	963	879
30	1048	"	880
20	1080	962	"
10	-	959	"
3	-	950	874
0	1170	1170	864

Lead sulfate ($PbSO_4$) + Lead tungstate ($PbWoO_4$)

Jaeger and Germs, 1921

mol%	f.t.	m.t.	tr.t. I	tr.t. II
100	1123	-	-	877
97	-	-	-	860
90	-	1070	-	"
85	1071	1042	876	859
80	1059	1011	"	858
75	-	985	"	"
70	1034	994	"	859
60	1010	995	875	"
50	-	"	"	858
45	1007	"	-	-
40	1020	"	874	859
30	1052	996	"	860
20	1086	995	873	859
10	1130	"	874	-
2	-	-	868	-
0	1170	-	864	-

Lead chromate ($PbCrO_4$) + Lead molybdate ($PbMoO_4$)

Jaeger and Germs, 1921

mol%	f.t.	tr.t. 1	tr.t. 2
0	844	783	707
5	837	798	706
10	"	799	705
15	"	"	704
20	838	"	703
25	"	"	700
30	"	-	697
40	839	800	"
50	"	"	"
60	872	-	-
70	924	-	-
80	969	-	-
90	1021	-	-
100	1065	-	-

Lead chromate ($PbCrO_4$) + Lead tungstate ($PbWoO_4$)

Jaeger and Germs, 1921

mol%	f.t.	mol%	f.t.
0	844	50	837
5	837	60	848
10	"	70	933
20	"	80	1016
30	"	90	1075
40	"	100	1123

Lead molybdate ($PbMoO_4$) + Lead tungstate ($PbWoO_4$)

Jaeger and Germs, 1921

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	1065	-	55	1096	1083
10	1070	1067	60	1099	"
20	1075	1070	70	1106	1091
30	1080	1074	80	1112	1100
40	1085	1079	90	1119	1111
45	1089	1082	100	1123	-
50	1093	1082			

Zinc chloride ($ZnCl_2$) + Zinc sulfate ($ZnSO_4$)

Evseiev and Bergman, 1951

mol%	f.t.	
0	322	A
2.6	320	"
4.8	318	"
7.0	310	"
9.6	300	E
13.8	340	(1+1)
17.6	356	"
22.7	322	"
27.2	410	"
32.5	436	E
38.0	490	B
43.6	556	"
49.0	600	"
54.0	627	"
59.0	658	"

Zinc potassium sulfate ($ZnK_2S_2O_8.6H_2O$) +
Zinc ammonium sulfate ($Zn(NH_4)_2S_2O_8.6H_2O$)

Krickmeyer, 1896

vol%	d 20°
100	1.937 (19°)
84.18	.967
77.68	.996
17.85	2.196
11.04	.200
12.00	.229
0	.252 (22°)

Cadmium fluoride (CdF_2) + Cadmium chloride ($CdCl_2$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
100	595	80	550
90	555	70	590
85	530	0	above 1000

Cadmium fluoride (CdF_2) + Cadmium iodide (CdI_2)

Ruff and Plato, 1903

mol %	f.t.	mol %	f.t.
100	340	80	345
95	335	75	355
90	335	0	above 1000

Cadmium chloride ($CdCl_2$) + Cadmium bromide ($CdBr_2$)

Greiner and Jellinek, 1933

mol% L	V	p_1	p_2
	752°		
0	0	53	0
30	42.7	54.5	39.6
50	58.6	41.9	59.4
70	80.5	29.1	121.2
100	100	0	200

Nacken, 1907

mol%	f.t.	mol%	f.t.
100	567	40	552
90	557	30	554
80	554	20	555.6
70	551	10	558
60	551	0	563
50	551.5		

Boardman, Palmer and Heymann, 1955 (fig.)

mol%	σ 600°	σ 700°	mol%	σ 600°	σ 700°
0	83.5	81	60	71	69
20	79.5	76.5	80	68.5	66
40	75	72.5	100	67	65

Boardman, Dorman and Heyman, 1949

mol %	molar volume (in cc)	mol %	molar volume (in cc)
	700°		
0	55.88	65.3	64.52
29.7	59.79	100	69.23
45.6	61.88		

Cadmium chloride ($CdCl_2$) + Cadmium iodide (CdI_2)

Nacken, 1907

mol %	f.t.	E	min.
100	385	-	-
90	376	360	108
80	367	358	210
70	360	360	378
60	384	359	300
50	412	360	258
40	440	358	198
30	473	359	138
20	502	359	90
10	532	359	48
0	563	-	-

Ilyasov and Bergman, 1956.

%	f.t.	%	f.t.
100	388	42.5	429
95	383	34.6	452
86	376	28	472
74.6	365	22.9	490
69	359E	16.1	513
62.9	375	5.3	544
52	399	0	568

Cadmium chloride ($CdCl_2$) + Cadmium sulfate ($CdSO_4$)

Ruff and Plato, 1903

mol%	f.t.	mol%	f.t.
0	590	20	575
10	570	30	635
15	565	100	1000

Bergman and Bakumskaya, 1955 and 1956

mol%	f.t.	mol%	f.t.
0	568	20	584
5	560	25	620
10	550	30	652
12.5	546	35	684
15	540	100	1000
17.5	564		

Semenkova, Bergman and Lesnykh, 1956

%	f.t.	%	f.t.
0	567	17.5	584
5	564	20	602
7.5	560	22.5	618
10	554	25	638
12	544	27.5	650
13	538	30	664
14	544	35	700
15	558	40	734
E : 14 %	538°		

Cadmium chloride ($CdCl_2$) + Cadmium molybdate ($CdMoO_4$)

Lesnykh, Bergman and Bikun, 1956 (fig.)

mol%	f.t.	mol%	f.t.
0	560	20	760
4.5	553 E	30	825
10	625		

Cadmium bromide ($CdBr_2$) + Cadmium iodide (CdI_2)

Nacken, 1907

mol%	f.t.	mol%	f.t.
0	567	60	429
10	539	70	410
20	520	80	396
30	493	90	389
40	472.5	100	385.5
50	451		

Hägg and Linden, 1943 (fig.)

mol %	spacing halogen layers (in Å)	mol %	spacing halogen layers (in Å)
0	3.13	60	3.35
20	3.22	80	3.40
40	3.29	100	3.43

Mercuric chloride ($HgCl_2$) + Mercuric bromide ($HgBr_2$)

Losana, 1926

%	f.t.	m.t.	tr.t.
0	275.6	-	-
5.09	270.5	264.6	-
11.46	265.2	257	-
20.18	258	249	-
25.40	255	247	-
32.15	251	244	210
40.22	246	240.8	215
45.00	243	238	218
50.18	239	236.8	219
53.02	237	235.6	220
55.00	235	234.2	221
54.86	234.3	231.5	220
65.00	232.5	228.4	218
69.80	230.1	226	216
75.18	227	224.2	-
77.89	225	223.5	-
80.92	223	222	-
84.78	225	222.4	-
90.16	228	224	-
92.15	229	225.5	-
95.00	231.5	227	-
100.00	235.6	-	-

Meerman and Scholten, 1939

mol%	tr.t.	mol%	tr.t.
53.0	5	62	74
58.0	53	64	107
59.5	66	67	121

van Nest, 1909

%	f.t.	d 20°	%	f.t.	d 20°
0	265	5.451	56.23	219	5.612
8.03	237	-	71.28	214	-
24.09	234	-	72.41	213	5.918
25.76	233	-	76.11	212	-
26.91	-	5.612	90.88	219	-
28.34	228	-	93.11	-	5.96
47.97	220	-	100	222	6.064

Mercuric chloride ($HgCl_2$) + Mercuric iodide (HgI_2)

Padoa and Tibaldi, 1903

mol%	f.t.	mol%	f.t.
100	254.0	36.4	180.3
94	242.7	33.4	188.9
88.4	230.5	32.3	194.7
78.8	210.1	30.6	198.8
72.5	199.2	24,6	214.5
66.2	185.7	21.4	222.9
60.7	174.1	17.9	232.9
55.7	163.5	14.3	241.2
50.9	153.3	9.8	252.7
47.2	147.8	7.2	258.9
44	150.0	1.8	270.6
41.2	161.2	0	277.4
38.7	172.4		

mol% C	mol% L
7.8	14.3
20.4	30.6
27.7	39.6
52.4	51.2
77.9	75.7

mol%	tr.t.	mol%	tr.t.
91.7	100	59.6	65
78.5	94	57.3	61
65.3	78		

mol% initial	mixed crystals sublimate
66.36	45.61
32.87	41.61

Bergman and Gänke, 1926

mol%	f.t.	E	tr.t.
0	281	-	-
2.5	276	145	-
5	269	"	-
10	255	"	-
20	230	"	-
33.3	188	"	-
50	151	"	74
55	161	"	73
66.7	191.5	"	"
75	210	"	"
82.5	221	-	80
90	237	-	101
100	257	-	127

Losana, 1926

%	f.t.	m.t.	E	min.	tr.t.
100	253	-	-	-	-
95.14	238.5	234	-	-	120.2
87.90	213	203.6	-	-	103
90.11	198	186	-	-	92
73.98	181.5	164	-	-	78
72.50	177.4	156	-	-	70
69.58	170	149	-	-	63
67.94	167.2	-	144	20	60
64.88	159	-	"	65	-
59.02	146	-	"	170	-
57.39	-	-	"	210	-
54.38	156.5	-	"	170	-
42.81	193.2	-	143.8	-	-
39.65	201	-	144	70	-
33.51	217	-	146	35	-
30.30	223.8	-	149	-	-
28.76	226.2	170	-	-	-
25	235	193	-	-	139.8
20	244	216	-	-	135.2
14.94	253.2	232	-	-	128
10	262	239	-	-	127
6.20	268	249	-	-	103
4.83	270	263	-	-	98
0	275.5	-	-	200	-

Bergman and Chagin, 1940

%	κ 200°	225°	250°	275°	300°
100	-	-	-	305	280
90	-	-	320	307	292
80	-	295	281	266	252
70	231	225	219	213	207
60	160	160	160	160	160
50	90	94	97	101	104
40	49	53	56	59	63
30	-	23	27	30	33
20	-	-	11	16	20
10	-	-	-	3	14
0	-	-	-	-	8

Mercuric chloride ($HgCl_2$) + Mercuric sulfate ($HgSO_4$)

Paic, 1930

mol%	f.t.
100	240
99	350
95	400
90	520
0	850

E : 1% 240°

Mercuric bromide ($HgBr_2$) + Mercuric iodide (HgI_2)

Reinders, 1899-1900

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	236.5	236	50	217.3	216
10	228.8	226	60	221.1	218
20	222.2	219	70	227.8	223
30	217.8	217	80	236.2	231
35	216.6	216	90	245.5	242
40	216.1	215.5	100	255.4	254
45	216.3	216			

tr.t.	% I red crystals	% II yellow crystals
127	100	100
100	98.5	94.8
80	97.4	90.5
50	95.7	84.5
40	95.0	92.5
25	93.8	78.7
0	91.4	77.0

% L	% C	% L	% C
70	71.1	35	33.4
60	61.5	30	29.5
50	50.9	20	18.8
40	40.1		

van Nest, 1909

%	f.t.	%	f.t.
0	222	51.70	207
9.98	205	69.97	210
18.87	205	70.71	213
19.21	-	76.58	212
23.66	207	77.65	214
33.15	205	80.09	215
41.05	205	100	245

Losana, 1926

%	f.t.	m.t.
94.85	246	243.4
84.98	237.4	231
70.58	225.6	220.3
47.82	217	216
31.70	219	217
23.78	223	219
10	231	227.1
3.84	234	232.2

Bruni and Padoa, 1912

Mixed crystals

% initial	% sublimate
19.92	18.31
43.31	36.52
60.06	50.44
80.02	71.74

van Nest, 1909

%	d	%	d
	20°		
0	6.064	51.70	6.179
9.98	-	69.97	6.283
18.87	-	70.71	6.311
19.21	6.092	76.58	-
23.66	6.152	77.65	-
33.15	6.165	80.09	-
41.05	6.159	100	6.303 and 6.282

Beck, 1907

mol%	f.t.
0	236.5
41	216.1
100	255.4

t	d	η	t	d	η
	0%			41 mol%	
240	5.25	3310	220	5.35	4594
247	5.23	2972	247	5.28	2923
257	5.21	1976	258	5.26	2550
	100%				
258	5.33	3547			

Mercuric bromide ($HgBr_2$) + Mercuric sulfate ($HgSO_4$)

Paic, 1930

mol%	f.t.	mol%	f.t.
0	235	20	470
5	350	30	510
10	395	100	850

E : 1% 236°

Mercuric iodide (HgI_2) + Mercuric sulfate ($HgSO_4$)

Páic, 1930

mol%	f.t.	E	mol%	f.t.	E
0	258	242	50	332	242
10	249	242	60	380	-
14	242	242 E	70	448	-
20	260	242	80	540	-
30	292	242	100	850	-
40	320	242			

(1+1)

Manganese oxide (MnO) + Manganese sulfide (MnS)

Andrew, Maddocks and Fowler, 1931

%	f.t.	E	%	f.t.	E
0	1585	-	50	1265-1285	-
5	1567	1530	60	1390	1285
10	1560	1280	65	1405	1285
15	1565	1270	70	1440	1255
20	1565	1285	75	1485	1280
25	1530	1270	80	1510	1280
30	1515	1260	90	1570	1550
35	1490	1280	95	1580	1560
40	1430	1280	100	1620	
45	1285	1255			

Manganese sulfide (MnS) + Manganese silicate ($MnSiO_3$)

Voloskoff, 1911

mol%	f.t.	E	mol%	f.t.	E
100	1216	-	77.88	1324	1132
96.57	1180	1125	69.88	1425	1134
93.15	1130	1170	60.17	1549	1128
87.07	1201	"			

Glaser, 1926

%	f.t.	E	%	f.t.	E
100	1274	-	86	1240	1090
99	1230	-	85	1240	1090
98	1220	-	84	1245	1080
97	1175	1080	83	1246	1075
96	1130	1080	82	1250	1070
95	1120	1070	80	1250	1070
94	1080	1080	75	1240	1080
93	1193	1070	70	1250	1070
92	1230	1080	50	1325	-
90	1215	1080	25	1410	-
88	1240	1095	0	1530	-

Andrew, Maddocks and Fowler, 1931

%	f.t.	E	%	f.t.	E
100	1305	-	65	1450	1255
99	1305	-	50	1530	1256
98.5	1305	1300	40	1550	1250
98	1300	1245	30	1575	1255
95	1285	1245	20	1575	1250
90	1250	-	10	1600	1250
80	1340	1230	0	1620	-
(heating curve)					
100	1257	-	80	1187	1117
98.5	1200	1100	65	1410	1115
95	1215	1172	50	1515	1115
90	1176	1150	40	1520	1115
85	1185	-	10	1520	1110
(freezing curve)					

Manganese silicate ($MnSiO_3$) + Manganese titanate ($MnTiO_3$)

Smolenski, 1912

mol%	f.t.	m.t.	E	tr.t.
0	1218	-	-	-
11.3	1211	-	1126	-
16.8	1175	-	1122	-
22.3	-	-	1120	-
33	1170	-	1121	1060
38.3	1189	-	1118	1060
43.4	1216	1155	-	1055
53.5	1256	1189	-	1053
63.3	1284	1250	-	1050
72.8	1311	1264	-	1048
82.1	1346	1310	-	1044
91.2	1370	-	-	-
100	1404	-	-	-

Ferrous chloride ($FeCl_2$) + Ferrous bromide ($FeBr_2$)

Mac Laren and Gregory, 1954 (fig.)

mol%	unit cell parameters (in Å)		
	a_o	c_o	
		HCP	CCP
100	3.80	18.6	-
80	3.75	18.55	-
60	3.71	18.4	-
50	3.69	18.3	18.1
40	3.675	18.25	18.0
30	3.66	18.15	17.9
20	3.64	-	17.8
0	3.61	-	17.6

Ferrous sulfide (FeS) + Ferrous orthosilicate (Fe_2SiO_4)

Andrew and Maddocks, 1932

%	f.t.	E	%	f.t.	E
0	1160	-	40	1020	-
2.5	1140	990	60	1020	-
5	1140	990	80	1062	1010
10	1125	970	85	1072	1000
20	1085	980	90	1095	-
		(heating curve)			
0	1150	-	50	1000	-
2.5	1150	980	60	995	930
5	1155	990	80	1025	920
10	1110	975	85	1055	-
20	1085	970	90	1030	-
40	1005	982			
		(freezing curve)			

Cobalt chloride ($CoCl_2$) + Cobalt sulfate ($CoSO_4$)

Lesnikh and Bergman, 1953

mol%	f.t.	mol%	f.t.
0	732	28	663
5	713	30	658 E
10	700	32	673
15	690	35	700
20	680	40	760
22	676	45	800
25	672		

Cobalt potassium sulfate ($K_2CoO_8S_2$) + Cobalt ammonium sulfate ($CoH_8N_2O_8S_2$)

Krickmeyer, 1896

Mixed crystals

vol%	d
	20°
100	1.905 (18°)
84.70	.928
27.39	2.148
14.85	.174
14.67	.178
0	.218

Nickel potassium sulfate ($K_2NiO_8S_2$) + Nickel ammonium sulfate ($NiH_8N_2O_8S_2$)

Krickmeyer, 1896

Mixed crystals

vol%	d
	20°
100	1.929
89.69	.951
88.19	.957
22.68	2.174
14.64	.196
14.14	.202
12.11	.213
0	.244

Sodium aluminum fluoride (Na_3AlF_6) + Potassium aluminum fluoride (K_3AlF_6)

Naray-Szabo and Sigmond, 1941

%	f.t.	%	f.t.
0	1001	49	941
10	986	50	941
15	974	55	943
20	964	56.74	942
25	953	60	941
30	946	65	951
32.5	941	67.5	953
33.5	941	70	950
35	942	71	951
36	939	73	948
37	941	75	951
38	939	77	951
39	936	80	945
40	938	85	958
45	943	90	968
46	943	100	991
47.5	941		

Aluminum chloride ($AlCl_3$) + Aluminum bromide ($AlBr_3$)

Pushin and Makucz, 1938

mol %	f.t.	E
0	190	-
24	150	70
29	136	71
35	121	73
40	107.5	74
42.5	98	75
49	88	75
52.5	-	75
58.5	75	75
69.5	79	76
73	81.5	77
75.5	85.5	78.5
83	89	83
91.5	93	89
100	97	-

Izbekov, 1925

mol%	f.t.	mol%	f.t.
100	97.4	59.4	84.3
95.1	92.3	53.6	96.7
91.0	88.5	48.3	106.6
83.7	83.1	42.4	119.8
78.8	79.0	35.0	133.8
73.2	76.7	23.9	149.8
65.2	73.1		

Aluminium oxide (Al_2O_3) + Aluminium carbide (Al_4C_3)

Baur and Brunner, 1934

%	f.t.	%	f.t.
0	2055	24.5	2048
6.75	2044	30	2082
13.5	2007	33.7	2098
18	2032	37.5	2118

Aluminium oxide (Al_2O_3) + Cryolithe (Na_3AlF_6)

Lorenz, Jabs and Eitel, 1913

%	f.t.	m.t.	E	min.
100	999	-	-	-
97.4	994	-	-	-
96.1	992	-	-	-
95.5	985	981	-	-
94.9	980	981	-	-
94.0	-	970	-	-
93.3	974	963	-	-
89.4	979	948	-	-
89.1	963	944	941	0
86.1	948	-	939	1.8
84.0	944	-	939	6.7
83.1	941	-	937	10.2
81.4	940	-	934	7.4
80.2	939	-	939	4.9
79.2	945	-	958	4.0
75.9	955	-	940	4.2
74.4	1001	-	941	0.7
71.2	1000	-	951	-
70.1	999	-	938	-
66.9	1072	-	939	9.7
0	above 2000	-	-	-

Fedotyev and Iljinski, 1911

mol%	wt%	f.t.	E	tr.t.
100	100	1000	-	565
95	97.5	985	965	"
90	95	975	945	555
80	89.2	950	925	"
75	86.1	940	930	"
72.5	84.5	935	930	"
70	82.8	945	935	560
65	79.3	1000	940	-

Sodium aluminum oxalate ($Na_6Al_2C_{12}O_{24}.9H_2O$) + Ammonium aluminum oxalate ($N_6H_{24}Al_2C_{12}O_{24}.6H_2O$)

Stortenbeker, 1913

Mixed crystals

%	d	angle of optical axes. 2 E	
0	1.871	83.301	isomorphe with sodium salt
3.1	.863	77.121	
5.6	.850	68.481	"
10.3	.840	70.31	"
13.7	.833	71.121	"
18.7	.821	74.31	"
44.6	.756	-	form I
49.1	.749	-	"
55.2	.737	-	"
69.2	.708	-	form II
100	.672	-	"

Ammonium alun ($AlH_4NS_2O_8$. 12 H_2O)
+ Potassium alun ($KAlS_2O_8$. 12 H_2O)

Krickmeyer, 1896

vol%	d
	20°
0	1.645
18.91	.662
36.90	.688
46.59	.697
46.85	.698
62.15	.717
75.40	.731
80.58	.732
87.74	.742
100	.757

Vanadium oxide (VO) + Vanadium nitride (VN)

Epelbaum and Brager, 1946

%	Unit cube edge (in Å)
0	4.08
14.7	4.076
44.7	4.095
62.9	4.100
68.5	4.106
68.0	4.110
81.6	4.116
82.7	4.120
88.1	4.124
90.4	4.126
96.0	4.128
100	4.129

Indium arsenide (InAs) + Indium antimonide (InSb)

Shih and Peretti, 1953

wt%	mol%	f.t.	E
0.0	0.0	942	-
12.5	10.1	926.3	-
20.0	16.8	917.5	524.8
40.0	34.7	875.1	525.3
60.0	54.3	844.6	526.2
80.0	76.6	760.6	524.8
90.0	87.9	684.7	527.3
92.0	90.5	670.5	524.9
95.0	94.0	632.0	525.8
96.0	95.1	618.3	526.3
97.5	97.2	595.2	524.2
100.0	100.0	525.0	-

Antimony trifluoride (SbF_3)
+ Antimony pentafluoride (SbF_5)

Lecat, 1949

%	b.t.
0	319
-	390 Az
100	155

Antimony trichloride ($SbCl_3$)
+ Antimony tribromide ($SbBr_3$)

Bernardis, 1912

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	73	-	60	56	55
10	68.5	64	70	55	54
20	64.5	58.5	80	63	58
30	62	56.5	90	75	67
40	59.5	55.5	100	93	-
50	57	55			

Antimony trichloride ($SbCl_3$) + Antimony triiodide (SbI_3)

Bernardis, 1912

mol%	f.t.	E	min.
0	73	-	-
5	59.5	41.5	20
10	50	41	40
20	42.5	41	150
30	62	41.5	100
40	78	41.5	40
50	89.5	40.5	20
60	103	-	-
70	116.5	-	-
80	129	-	-
90	145	-	-
100	165	-	-

Antimony tribromide ($SbBr_3$) + Antimony triiodide (SbI_3)

Bernardis, 1912

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	93	-	50	118	105
5	89.5	86	60	127.5	118
10	86	83	70	137	130
15	84.5	82	80	148	143
20	88	84	90	156.5	154
30	96	87	100	165	-
40	105.5	96			

Antimony trioxide (Sb_2O_3) + Antimony trisulfide (Sb_2S_3)

Quereigh, 1912

mol%	f.t.	E	min.	tr.t.	min.
0	656	-	-	-	-
2.5	644	-	-	-	-
5	632	-	-	-	-
10	616	-	-	-	-
15	608	-	-	-	-
20	592	487	30	-	-
30	580	486	70	-	-
40	569	489	140	-	-
50	560	488	150	-	-
60	526	489	170	-	-
66.6	-	488	190	-	-
70	500	489	120	-	-
75	513	488	30	-	-
77.5	519	487	-	-	-
80	-	-	-	522	70
83.3	530	-	-	521	110
85	530	-	-	522	110
90	544	-	-	522	90
95	547	-	-	522	60
100	548	-	-	-	-

(1+1)

Bismuth fluoride (BiF_3) + Bismuth trioxide (Bi_2O_3)

Aurivillius, 1955

%	phases	%	phases
78.9	BiF_3	83.3	BiOF + III
80	BiF_3 + I	84.8	III
80.3	I	85.5	IV
82.7	I + II	87.2	IV
85.2	II	100	Bi_2O_3
87.5	II		

I, II, III and IV : complexes of unknown compositions .

Titanium oxide (TiO) + Titanium nitride (TiN)

Schmitz-Dumont and Steinberg, 1954 (fig.)

mol%	d	mol%	d
0	4.90	50	5.11
20	4.83	60	5.30
30	4.80	80	5.30
40	4.97	100	5.30

mol%	lattice constant Å
0	4.170
20	.192
30	.208
40	.228
60	.248
80	.248
100	.248

Bismuth triselenide (Bi_2Se_3) + Bismuth tritelluride (Bi_2Te_3)

Amadori, 1915

%	f.t.	E	min.
100	575	-	-
98.06	-	570	370
96.07	-	"	350
93.35	584	"	290
87.17	598	"	200
78.43	610	"	130
75.70	612	"	70
70.04	614	"	50
65.57	615	"	20
60.90	615	-	-
50.96	642	614	170
40.04	670	"	120
35.18	685	"	90
21.57	712	612	50
0	-	-	-

(1+1)

Stannic chloride ($SnCl_4$) + Stannic bromide ($SnBr_4$)

Räder, 1923

mol%	b.t.	f.t.	m.t.
0	+114.0	-33.0	-
5.0	-	-36.0	-37.0
10.2	-	-37.0	-37.8
17.0	+126.0	-37.0	-37.7
35.5	+139.5	-30.0	-32.3
50.0	+153	-19.5	-24.5
60.3	+161.5	-10.3	-18.5
70.0	+169	-1.6	-11.6
80.2	+179	+8.2	-0.3
90.2	+197	+18.0	+15.0
100	+206.7	+29.2	-

Stannic chloride ($SnCl_4$) + Stannic iodide (SnI_4)

Räder, 1923

mol%	tr.t. I	tr.t. II	f.t.
0	-	-	-33.0
5.7	-61.0	-	-42.5
10.0	-60.5	-	-50.0
15.2	-60.0	-	-
20.0	-61.0	-	-54.0
25.0	-60.5	-	-50.7
30.2	-60.3	-	-47.3
32.6	-60.2	-46.5	-43.5
34.5	-	-46.5	-35.0
40.3	-	-47.0	-15.5
50.0	-	-47.0	+14.0
59.5	-	-50.0	+42
69.8	-	-50.0	+70
80.0	-	-	+94
90.0	-	-	+120
100	-	-	+145.3

Stannic bromide ($SnBr_4$) + Stannic iodide (SnI_4)

Auger, 1909

mol%	f.t.	mol%	f.t.
0	32	40	65
18	19	70	75
30	35	80	105

Räder, 1923

mol%	b.t.	mol%	b.t.
0	206.7	56.5	247
6.5	208	63.0	259
13.3	211.5	69.0	267
28.6	220.8	74.8	277
39.7	228.0	86.0	302.5
43.6	233.5	94.5	324.5
47.0	238.0	100.0	346.0

mol%	f.t.	mol%	f.t.
0	29.2	50	54.8
10	22.4	56.6	66
20	19.5	63.3	77.4
21	19.4	69.0	87.8
22	19.35	75.0	98.2
23	19.35	86.0	122.1
24	19.45	92.5	134.5
25	20.05	94.6	138.0
31.6	29.15	100.0	145.3
41.6	41.8		

% L	C	% L	C
10	8.7	50	61.5
22.5	22.5	72.1	80.7
33.5	40.5		

0. WATER + SALTS OF MONOVALENT METALS

XLIV. WATER + SODIUM SALTS .

Water (H_2O) + Sodium fluoride (NaF)

Robinson, 1941

Isopiestic solutions	m_1	m_2	m_1	m_2
		25°		
	0.1554	0.1562	0.6765	0.6896
	0.2490	0.2504	0.7874	0.8088
	0.4047	0.4097	0.8518	0.8734
	0.5616	0.5728	0.8908	0.9124
	0.5774	0.5870	0.9106	0.9356

H_2O + Sodium chloride (NaCl)

Heterogeneous equilibria .

Wüllner, 1858

Vapour pressure

t	p 0%	4.8%	9.1%	13.0%	16.7%	20.0%	23.1%
19.9	17.28	16.30	15.81	15.01	-	-	13.22
24.2	22.45	21.27	20.37	19.48	-	-	17.29
29.9	31.36	30.25	29.31	27.88	-	-	24.97
30.8	33.02	-	30.14	28.85	27.75	26.06	-
35.0	41.82	40.03	38.71	37.12	-	-	33.69
35.4	42.74	-	39.40	37.97	34.73	34.99	-
39.5	53.43	-	49.66	47.97	46.28	44.10	-
40.9	57.61	55.54	54.05	51.86	50.13	-	46.00
42.6	62.97	-	58.81	57.02	54.64	51.96	-
44.8	70.27	68.08	66.15	61.47	-	-	59.18
45.7	73.98	-	69.11	67.13	64.56	61.58	-
48.4	84.36	81.49	79.20	76.43	72.74	-	66.52
49.1	87.93	-	82.29	79.81	76.44	73.07	-
49.8	91.09	88.25	85.28	82.11	79.73	-	72.89
52.5	104.49	101.41	98.34	95.82	92.05	-	83.57
53.8	110.87	-	103.84	101.27	96.62	93.15	-
54.1	112.59	109.12	104.95	101.98	99.03	-	90.41
56.6	126.84	-	118.33	115.15	110.69	106.65	-
57.9	134.88	130.83	126.77	123.80	118.26	-	108.04
59.0	142.01	137.89	134.03	129.48	125.69	-	-
60.6	152.99	-	142.61	139.22	-	130.04	-
60.9	155.09	-	145.00	141.19	136.30	131.55	-
61.2	157.29	152.54	-	143.83	137.60	-	-
62.2	165.07	-	154.80	-	144.56	139.17	-
63.8	177.11	-	167.18	162.33	156.35	149.73	-
64.7	184.45	-	173.28	169.54	-	158.60	-
64.8	185.27	-	175.39	170.25	163.82	157.30	-
67.8	211.73	-	198.79	193.95	-	181.66	-
67.9	212.67	-	200.82	194.60	186.23	182.45	-
68.6	219.36	212.53	206.52	201.56	191.87	-	-
68.7	220.24	-	208.20	201.93	194.75	-	-
72.2	256.27	-	241.76	234.85	226.07	216.49	-
75.3	292.20	-	275.73	-	258.00	247.06	-
75.5	294.66	-	277.69	-	258.65	248.00	-
78.6	335.01	-	-	-	296.07	285.72	-
82.2	388.33	-	366.75	-	342.13	328.92	-
85.7	445.09	-	419.79	-	392.51	379.19	-
88.5	496.15	-	467.51	-	436.21	-	-
91.2	549.92	534.08	517.55	-	483.22	-	-
92.0	566.76	-	532.93	-	497.21	-	-
92.0	566.76	-	532.52	514.92	492.00	-	-
100.5	775.40	-	730.50	712.48	682.99	-	-
100.7	779.26	-	729.95	-	684.27	-	-

Pauchon, 1880

t	p	t	p	t	p
		21.30%			
7.89	7.04	18.87	14.09	32.85	31.40
8.11	7.14	19.29	14.57	33.31	32.47
8.35	7.29	19.99	15.21	33.78	32.93
8.75	7.46	20.45	15.56	34.57	34.21
9.18	7.71	20.88	15.98	34.99	35.07
9.77	7.97	21.37	16.45	35.40	35.79
10.63	8.44	21.95	17.05	35.90	36.71
10.99	8.68	22.31	17.43	36.50	37.93
11.91	9.24	22.67	17.86	37.01	38.98
12.31	9.35	23.01	18.27	38.76	42.43
13.17	9.98	24.07	18.27	39.55	44.14
13.65	10.27	25.12	20.40	40.15	45.42
14.43	10.83	25.83	21.32	40.65	46.51
14.81	11.08	26.52	22.19	41.33	47.37
15.22	11.29	26.92	22.68	42.01	49.49
16.01	11.94	27.32	23.23	43.56	51.25
16.83	12.51	28.85	26.27	43.60	53.35
17.32	12.88	29.30	25.95	43.61	53.57
17.98	13.47	29.87	26.69	44.18	55.07
18.26	13.65	30.39	27.51		
18.46	13.79	30.92	28.31		
18.67	14.08	31.45	29.12		
		31.98	29.79		
		13.15 %			
7.52	7.33	15.85	12.49	27.99	26.21
7.83	7.47	16.62	13.17	28.50	26.04
8.41	7.76	17.43	13.88	28.75	27.23
8.97	8.03	18.29	14.59	28.73	27.28
9.51	8.35	19.47	15.76	29.17	28.02
9.98	8.55	20.08	16.36	29.88	29.24
10.41	8.87	20.77	17.01	30.62	30.45
10.85	9.19	21.19	16.59	30.64	31.41
11.28	9.41	21.88	18.25	31.18	32.65
11.75	9.65	22.46	18.87	31.87	33.51
12.15	9.91	22.96	19.52	32.32	34.64
12.18	9.88	23.54	20.21	32.94	35.59
13.16	10.57	24.11	20.73	33.41	36.51
13.17	10.57	25.73	22.97	33.91	37.79
14.05	11.21	26.37	23.78	34.51	38.88
14.65	11.55	27.02	24.71	35.01	
15.03	11.94	27.64	25.66		

Müller-Erzbach, 1885

%	t	D_p	%	t	D_p
7.29	11.33	0.49	13.12	13.60	1.20
7.88	13.5	.46	14.93	13.80	1.35
8.69	12	.49	17.28	14.25	1.70
9.10	8.80	.53	20.18	13.89	1.95
9.52	6.00	.53	24.11	13.20	2.73
10.62	12.20	.82	26.37	13.88	3.05
11.83	12.50	.97			

$D_p = pH_2O - psol$

Tammann, 1885

t	p 0%	p 12.87%	p 17.99%	p 26.29%
49.13	88.6	80.7	76.0	66.7
52.21	103.2	93.3	88.7	78.2
55.58	121.4	110.9	104.9	92.0
58.90	161.1	129.8	120.3	108.4
61.63	177.4	146.8	139.3	121.6
63.75	194.8	162.0	141.1	134.8
65.83	218.9	177.5	167.8	148.2
68.49	245.9	199.0	188.5	166.2
71.17	285.9	224.7	211.3	186.8
74.72	292.5	260.2	246.1	217.6
75.27	323.2	266.8	252	222.2
77.66	417.1	294.6	278.9	246.5
84.00	431.4	381.1	360.3	319.5
84.86	472.3	398.5	371.7	329.0
87.19	503.3	330.6	307.1	359.9
88.84	541.5	456.0	433.8	384.2
90.77	598.1	494.6	467.7	413.5
93.42	670.4	546.7	516.9	457.9
96.53	770.9	611.9	578.7	512.6
100.40	142.5	703.7	666.6	590.1

%	p	%	p
100°			
0	760.0	18.41	652.9
5.50	734.9	21.87	624.4
10.35	709.7	26.96	572.5
14.57	681.3		

Nicol, 1886

mol %	D_p 70°	75°	80°	85°	90°	95°
2	8.5	10.8	13.7	16.4	20.4	24.6
3.9	18.0	22.5	28.0	33.8	40.9	50.2
4.8	22.6	28.5	35.2	43.0	52.5	63.3
5.7	28.4	35.4	43.5	52.5	63.5	77.0
7.4	39.0	48.6	59.9	72.8	88.2	107.1
9.1	50.4	62.4	76.6	93.1	112.8	136.4

mol %	p 70°	75°	80°	85°	90°	95°
0	228.3	283.2	349.4	428.3	521.4	631.0
6.10	219.8	272.4	335.7	411.9	501.0	606.4
11.50	210.3	260.7	321.4	394.5	480.5	580.8
13.96	205.7	257.7	314.2	385.3	468.9	567.7
16.30	199.9	247.8	305.9	375.8	457.9	554.0
20.63	189.3	237.6	289.5	355.5	433.2	523.9
24.51	177.9	220.8	273.8	335.2	408.6	494.6

Emden, 1887

t	p	t	p	t	p
4.82 %		9.17 %		12.82 %	
18.92	15.8	18.11	14.65	13.99	10.9
25.44	23.4	25.48	22.9	23.93	20.3
30.13	31.8	29.58	29.15	28.38	26.1
36.56	44.3	35.25	40.00	35.01	38.0
40.88	55.55	40.87	54.1	38.83	46.8
44.92	69.5	43.63	62.5	45.94	68.3
49.69	87.3	50.36	87.0	49.19	79.9
55.09	113.6	54.71	108.45	55.23	107.8
60.09	144.7	60.43	141.8	57.20	118.5
64.58	177.9	64.52	170.2	70.38	215.15
70.85	234.3	70.15	218.3	75.73	269.3
75.86	289.7	74.55	264.4	80.45	325.7
82.33	367.2	80.68	340.4	85.53	400.4
85.99	434.41	90.87	506.8	90.66	487.05
91.36	534.2			95.19	577.8
16.72 %		21.05 %		23.13 %	
19.14	14.35	21.03	15.35	19.45	13.4
25.54	21.3	25.59	20.1	25.91	19.7
32.07	31.05	30.16	26.3	30.56	26.0
35.28	37.1	35.61	35.5	35.52	34.2
41.24	51.25	39.78	44.5	39.82	43.2
44.69	61.5	45.92	61.5	44.93	56.3
50.00	79.6	49.42	73.25	50.64	74.9
54.49	99.4	55.35	97.5	54.38	90.2
59.99	129.7	60.16	122.25	60.04	117.2
64.43	158.1	64.00	146.3	63.52	138.7
69.37	196.7	69.61	187.6	71.78	199.3
74.59	246.2	76.00	246.3	74.39	222.9
79.50	300.3	79.29	281.7	79.64	277.65
84.12	361.8	85.40	359.45	85.48	349.0
89.72	449.7	91.09	448.7	90.37	422.45
95.12	553.4	95.47	526.8	95.14	504.75

Dieterici, 1891

%	p	%	p
0°			
0	4.620	18.96	3.930
5.53	4.460	22.63	3.722
10.47	4.301	26.20	3.504
14.89	4.125		

Kahlbaum, 1893

p	% 0°	17.4°	35.4°
15	12.79	19.60	22.80
20	17.54	24.15	27.46
25	23.76	27.78	30.90
30	-	30.85	34.02
35	-	33.63	36.92
40	-	36.27	39.60

Smits, 1899 - 1900

m	D_p	m	D_p
	0°		
0.02842	0.00344	0.35587	0.05026
0.03546	0.00477	0.8854	0.12757
0.08813	0.01223	1.7533	0.25724
0.033028	0.00435	1.8228	0.26757
0.17680	0.02477	2.1927	0.33406
0.34057	0.04793	4.6362	0.78345
0.05185	0.00675	1.0307	0.14626
0.10733	0.01476	1.6078	0.23082
0.25770	0.03650		

Krauskopf, 1910

t	p	t	p
	0%		
24.9	23.49	60	149.26
25	23.711	60.2	150.65
30	31.84	70	233.99
35	41.91	70.05	234.46
40	55.18	80	355.08
50	92.16		

%	p	%	p
40°		60°	
4.85	55.67	7.48	144.02
9.42	51.74	11.79	137.94
13.71	49.36	16.38	132.09
18.52	47.36	18.52	125.38

Leopold and Johnston, 1927

t	m	p_1	p
20.42	6.130	18.19	13.65
25.49	.148	24.45	18.35
29.96	.167	31.75	23.80
36.92	.201	46.86	35.03
40.55	.220	56.96	42.54
50.00	.277	92.51	68.84

Ebert, 1930

%	p 0°	5.2°	18°	20°	30°
0	4.58	6.64	15.48	17.54	31.82
5	4.4	6.5	15.0	17.0	30.7
10	4.3	6.2	14.4	16.3	29.5
15	4.2	5.9	13.7	15.5	28.2
20	3.8	5.6	12.9	14.5	26.4
25	3.5	5.1	12.0	13.4	24.6
26.5	(satd)	-	-	-	23.9

Pearce and Nelson, 1932

M	p	M	p
	25°		
0.0	23.752	3.0	21.227
0.2	23.594	3.5	20.733
0.4	23.444	4.0	20.210
0.6	23.286	4.5	19.677
0.8	23.127	5.0	19.135
1.0	22.969	5.5	18.576
1.5	22.560	6.0	17.990
2.0	22.134	6.138	17.825
2.5	21.692		

Lannung, 1934

m	p	m	p
	18°		
0.633	15.17	2.226	14.30
0.641	15.14	2.661	15.05
0.729	15.13	2.919	13.95
1.021	15.01	3.483	13.52
1.882	14.52	3.941	12.23
2.010	14.44	5.281	12.28
2.153	14.32	5.437	12.18

Smith, Combs and Googin, 1954

m	p°-p/p°	m	p°-p/p°
	30.01°		
3.755	0.139	5.070	0.196
4.563	.174	5.162	.200
4.910	.190	4.910 (25.00°)	.188

van't Hoff, Armstrong and al., 1903

t	p	t	p
	sat. sol.		
25	17.7	50	68.7
30	24	55	86.4
35	31.6	60	109.5
40	42.1	65	138.2
45	54.3	70	173.4

Speranski, 1909 and 1910

t	p	t	p
23.88	16.79	24.15	17.02
27.48	20.72	27.48	20.72
31.1	25.43	37.2	35.72
34.4	30.61	43.89	51.0
37.2	35.77	47.24	60.59
40.32	42.29		
42.25	46.79		
45.32	54.92	(sat.sol.)	

Speranski, 1913

t	p	t	p
sat.sol.			
95.03	468.17	79.47	259.48
88.73	373.03	74.77	211.53
85.24	323.20	68.98	165.52
83.59	304.24	65.65	142.75
80.18	265.36	61.76	120.04
79.58	260.04	53.87	83.08
		47.4	59.58

Badger and Baker, 1920

t	p	t	p	t	p
sat. sol.					
107.90	735.7	86.24	329.9	65.86	139.0
107.80	731.6	85.13	319.5	65.70	138.3
103.84	575.0	83.66	301.5	63.85	127.0
99.90	558.0	82.44	281.9	63.36	123.9
97.92	518.7	82.25	280.3	62.73	120.2
94.68	456.9	78.56	242.5	61.82	116.7
94.66	458.0	78.16	239.3	60.27	109.6
93.83	444.9	77.87	236.7	60.05	108.4
93.39	437.3	75.30	213.5	57.38	92.9
91.37	406.0	75.23	212.1	54.79	84.2
90.38	391.8	73.98	201.5	51.54	70.0
90.02	381.5	68.11	155.3	50.61	67.7
89.98	381.0	68.04	154.0	50.04	64.8
86.36	331.6				

Pohle, 1927

t	p	t	p
sat. sol.			
30.0	26.0	70.0	175.0
35.0	32.0	75.0	216.0
40.0	42.0	80.0	266.0
45.0	54.0	85.0	327.0
50.0	71.0	90.0	392.0
55.0	90.0	95.0	471.0
60.0	113.0	100.0	566.0
65.0	141.0	105.0	676.0

Hepburn, 1928

t	p
sat. sol.	
20	13.34
25	17.75
30	23.90

Adams and Merz, 1929

t	p	t	p
sat. sol.			
10	7.00	30	23.96
15	9.87	40	41.37
20	13.63	50	68.50
25	18.01		

Bovalini, 1931

%	t	sat.sol.
20.369	30.4	32.278

Foote, Saxton and Dixon, 1932

Saturated solutions

$$\lg P = -\frac{2890.7}{T} - 4.715 \lg T + 22.612$$

Olander and Liander, 1950

Critical curve

%(V)	t	P Kg	%(V)	t	P Kg
0	374	224	3.0	408.5	309
0.2	379.5	239	4.0	416.5	332
0.5	383.5	248	5.0	424	354
1.0	388.5	259	6.0	431	376
1.5	394	271	7.0	437.5	397
2.0	399	283			

Morey and Chen, 1956

t	P kg
V+L+C	
374	139
400	176
500	337
600	414
700	403

Boiling point

Legrand, 1835

%	b.t.	%	b.t.	%	b.t.
sat. sol.					
0	100.0	15.47	103.0	24.13	106.0
4.21	100.5	17.15	103.5	25.31	106.5
7.15	101.0	18.77	104.0	26.36	107.0
9.75	101.5	20.32	104.5	27.38	107.5
11.81	102.0	21.69	105.0	28.42	108.0
13.71	102.5	22.96	105.5	29.18	108.4

Kremers, 1856

Sat. sol. b.t. = 109°

Guthrie, 1875

%	b.t.	%	b.t.	%	b.t.
26.27	108.4	11.82	102.2	3.94	-
18.39	104.3	10.51	102.0	2.627	-
17.07	103.8	9.19	101.6	1.97	-
15.76	103.6	7.88	101.3	1.31	-
14.45	103.0°	6.57	100.8	0.66	-
13.13	102.6	5.25	100.6	0.00	100.0

Baroni, 1893

%	b.t.	%	b.t.	%	b.t.
1.22	100.193	5.87	101.022	12.82	102.630
2.44	100.396	8.04	101.457	14.88	103.224
3.84	100.640	10.29	101.956		

Richmond, 1893

%	b.t.	%	b.t.	%	b.t.
7.6	102.2	16.1	104.8	24.0	107.7
11.0	103.0	18.8	106.1	26.0	108.7
14.9	104.2	22.3	107.1	28.7	109.5

Buchanan, 1899

p	b.t.	p	b.t.	p	b.t.
sat.sol.					
790	109.64	700	106.10	620	102.61
780	109.27	690	105.68	610	102.15
770	108.88	680	105.29	600	101.68
760	108.50	670	104.83	590	101.21
750	108.11	660	104.39	580	101.73
740	107.72	650	103.96	570	100.24
730	107.32	640	103.51	560	99.75
720	106.92	630	103.07	550	99.24
710	106.51				

%	b.t.	%	b.t.	%	b.t.
550.4 mm					
28.51	99.31	23.60	97.29	14.93	94.25
28.05	99.10	21.33	96.27	0.00	91.20
26.15	98.30	18.31	95.26		
555.6 mm					
-	99.58	26.25	98.57	0.00	91.47
29.01	99.53	23.32	97.56		
613.8 mm					
28.49	102.32	20.89	99.06	14.49	97.05
28.19	102.20	19.48	98.57	12.57	96.55
25.79	101.00	17.95	98.06	0.00	94.13
23.42	100.07	16.26	97.56		
619.2 mm					
29.00	102.58	24.00	100.51	18.48	98.49
28.77	102.41	21.37	99.49	0.00	94.35
26.28	101.51				
619.4 mm					
28.60	102.58	23.57	100.49	15.03	97.55
28.16	102.40	21.16	99.47	11.06	96.44
25.91	101.50	18.27	98.46	0.00	94.37
738.8 mm					
28.16	108.25	20.71	104.37	14.77	102.35
27.86	107.52	19.40	103.86	12.97	101.84
25.38	106.40	17.97	103.36	0.00	99.21
23.12	105.39	16.43	102.85		
753.3 mm (non satd)					
21.47	105.5	7.50	101.19	3.62	100.41
16.31	103.54	6.48	100.91	3.08	100.31
12.38	102.33	5.38	100.71	2.51	100.21
10.59	101.82	4.27	100.51	0.00	99.75
8.45	101.32				
761.6 mm					
28.09	108.54	24.91	106.96	21.29	105.49
27.92	108.44	23.71	106.47	20.10	104.99
25.91	107.55	22.59	105.98	0.00	100.06
772.0 mm					
28.16	108.98	21.58	105.99	16.23	104.02
27.78	108.80	20.38	105.80	14.62	103.52
26.00	107.97	19.12	105.01	12.72	103.03
23.93	106.88	17.75	104.51	0.00	100.44
22.67	106.43				

Kahlenberg, 1903

%	b.t.	%	b.t.	%	b.t.
747.0 mm					
1.13	100.20	7.63	101.57	13.72	103.35
4.00	100.75	9.92	102.19	17.02	104.56
5.23	101.02				
743.5 mm					
2.82	100.51	11.20	102.57	15.63	104.12
5.66	101.11	11.42	102.23	16.00	103.70
6.63	101.73	13.52	102.77	16.75	104.57
9.17	101.68	13.52	103.30		
754.0 mm					
1.24	100.60	4.80	100.83	8.09	101.49
2.43	100.39	5.93	101.04	10.23	101.99
3.66	100.60				

Johnston, 1906 and 1907

%	b.t.	%	b.t.
0.48	100.096	6.47	101.144
0.64	100.097	7.20	101.332
0.91	100.169	7.94	101.482
0.92	100.198	8.20	101.619
1.23	100.228	9.66	101.842
1.66	100.294	10.35	101.971
2.01	100.356	10.88	102.132
2.38	100.421	11.76	102.313
2.98	100.541	12.78	102.551
3.88	100.680	13.30	102.689
4.80	100.826	13.81	102.818
5.67	100.986	14.40	102.936

Berkeley and Applebey, 1911

Sat. sol. b.t. = 108.668°

Jablezynski and Kon, 1923

m	b.t.	m	b.t.
0.2085	100.191	1.5045	101.493
.4924	100.469	.7672	101.780
.9208	100.888	.9824	102.022
1.2250	101.200	2.2639	102.347

Gerlach, 1926

%	b.t.	%	b.t.
0	100	20.32	105
3.29	100.5	21.57	105.5
6.19	101	22.78	106
8.76	101.5	23.96	106.5
11.03	102	25.09	107
12.96	102.5	26.20	107.5
14.67	103	27.28	108
16.25	103.5	28.32	108.5
17.69	104	28.93	108.8
19.03	104.5		

%	b.t.	%	b.t.
0.8	100.1	10.3	101.8
1.6	100.3	11.5	102.3
3.1	100.4	14.0	102.8
3.2	100.5	18.8	104.4
4.9	100.7	24.3	106.6
6.1	101.0	24.5	106.8
9.8	101.7	26.0	107.4

Monrad and Bagder, 1929

p	b.t.	p	b.t.
	sat.sol.		
94.9	56.77	510.4	97.37
224.4	76.18	613.3	102.57
315.9	84.48	676.7	105.32
398.1	90.68	708.7	106.68

Styrikovich and Khaibullin, 1956 (fig.)

%	b.t.	%	b.t.
1 kg		5 kg	
0	100	0	160
20	112	20	170
28	120	30	175
10 kg		20 kg	
0	180	0	230
20	190	20	230
32	200	34	240
30 kg		40 kg	
0	240	0	260
20	250	20	270
35	270	38	280
50 kg		60 kg	
0	270	0	270
20	290	20	300
39	300	40	310
70 kg		80 kg	
20	300	0	300
40	320	20	310
		40	340
90 kg		100 kg	
0	310	20	340
20	320	41	350
40	350		
110 kg		120 kg	
20	340	20	340
42	360	43	370
140 kg		160 kg	
20	370	20	380
45	390	46	390
180 kg		200 kg	
20	380	20	390
47	400	48	410
220 kg		240 kg	
20	400	20	400
49	420	50	440
260 kg		300 kg	
20	430	20	430
52	460	55	470
320 kg		340 kg	
20	440	20	450
57	490	59	510
360 kg		380 kg	
20	450	20	450
62	550	63	560
400 kg (crit.p.)			
100	804		
100	2300 cr.t.		

mg/kg (in V)	t	mg/kg (in V)	t
30 kg/cm2		60 kg/cm2	
0.001	250	0.001	280
0.01	280	0.1	280
0.1	350	0.7	400
1	460	6	550
80 kg/cm2		100 kg/cm2	
0.001	300	0.001	310
1	300	8	310
1.5	400	6	450
8	550	9	550
120 kg/cm2		140 kg/cm2	
0.001	330	0.001	340
10	330	30	340
20	330	70	370
8	450	10	450
10	550	10	550
160 kg/cm2		180 kg/cm2	
0.001	350	0.001	360
80	370	200	380
20	480	50	450
20	550	50	550
200 kg/cm2		225 kg/cm2	
0.001	370	0.001	380
200	380	600	380
70	450	100	450

g/kg (in L)	t	g/kg (in L)	t
240 kg/cm2		260 kg/cm2	
0.1	500	0.4	500
5	380	10	290
100	380	800	440
800	430		
280 kg/cm2		300 kg/cm2	
0.6	480	0.8	480
10	400	10	410
800	460	800	440
320 kg/cm2		340 kg/cm2	
1	500	0.9	510
10	450	10	440
800	440	800	490
360 kg/cm2		380 kg/cm2	
0.9	530	1	550
10	450	10	440
800	470	800	520
400 kg/cm2		400 kg/cm2	
1	450	600	480

Freezing point .

Poggiale, 1843

%	f.t.	%	f.t.
24.66	-15	40	26.82
25.08	-10	50	27.00
25.49	-5	60	27.14
26.21	0	70	27.47
26.27	5	80	27.69
26.33	9	90	27.99
26.40	14	100	28.37
26.54	25	109.7	28.75

Rüdorff, 1861

%	f.t.	%	f.t.
0.99	-0.6	12.28	-8.4
1.96	-1.2	13.04	-9.2
3.85	-2.4	13.79	-9.9
5.66	-3.6	14.53	-10.6
7.41	-4.8	15.25	-11.4
9.09	-6.0	15.97	-12.1
10.71	-7.2	16.67	-12.8

Mulder, 1866

%	f.t.
26.31	0
26.52	25.5
26.58	44.5

Nordenskjold, 1869

%	f.t.
25.15	1.5
26.36	13.75
27.59	70.0
28.26	108.5

de Coppet, 1872

%	f.t.	%	f.t.
4.76	-2.9	16.67	-13.6
9.09	-6.1	17.94	-15.05
13.04	-9.7	22.58	-21.4
14.89	-11.45	23.6	-23.6

Guthrie, 1875

%	f.t.	%	f.t.
26.27	-	7.88	-5.4
18.39	-15.4	6.57	-4.1
17.07	-15.0	5.25	-3.4
15.76	-12.4	2.627	-2.1
14.45	-11.1	1.97	-1.5
13.13	-9.4	1.31	-1.9
11.82	-7.7	0.66	-1.5
10.51	-6.7	0.0	0.0
9.19	-5.7		

Guthrie, 1876

%	f.t.	%	f.t.
1	-0.3	19	-15.5
2	-0.9	20	-17.0
4	-1.5	22	-20.0
7	-2.2	23.6	-22.0
10	-4.2	25	-12.0
13	-6.6	26.27	0.0
15	-9.1	26.5	+25.0
16	-11.0	26.8	+26.8
16	-11.9		

De Coppet, 1883

%	f.t.	%	f.t.
26.69	-14.0	26.75	38.55
24.525	-14.0	26.83	44.75
24.33	-13.8	27.03	52.5
25.49	-6.25	27.01	55.0
25.37	-5.95	27.96	59.75
26.25	0.0	27.52	71.3
26.36	+3.6	27.52	74.45
26.44	5.3	27.75	82.05
26.44	14.45	27.78	86.7
26.26	20.85	28.01	93.65
26.43	25.45	28.95	101.7

Tilden and Shentone, 1884

%	f.t.	%	f.t.
28.47	118	30.36	160
29.63	140	30.98	

Andreae, 1884

%	f.t.	%	f.t.
26.28	0	26.83	50
26.28	4	27.04	60
26.31	10	27.28	70
26.38	20	27.54	80
26.48	30		
26.65	40		

Raupenstrauch, 1885

%	f.t.	%	f.t.
26.24	0	26.85	50
26.30	10	27.06	60
26.39	20	27.30	70
26.51	30	27.56	80
26.66	40		

Etard, 1894

%	f.t.	%	f.t.
25.5	-7	28.0	77
23.3	-17	28.2	90
23.5	-18	29.1	115
23.4	-21	28.9	135
23.4	-21	28.8	140
23.7	-21	29.6	150
25.8	0	30.2	180
26.7	+15	31.6	215
26.8	55		

Berkeley, 1904

%	f.t.	%	f.t.
26.34	0.35	27.16	61.70
26.40	15.20	27.44	75.65
26.55	30.05	27.81	90.50
26.79	45.40	28.39	107.0 (b.t.)

Jones and Getman, 1904

M	f.t.
0.2	-
0.5	-1.759
1.0	-3.546
2.0	-7.467
3.0	-12.223

Matignon, 1909

%	f.t.	%	f.t.
11	-6.6	25	-16.7
15	-9.25	30.7	-21.3 E
20	-12.7	32.9	+12.25
		(NaCl.2aq.)	

Schreinemakers, 1910-11

mol%	f.t.
10.00	35
9.95	25
9.90	15

Sudhans, 1914

mol%	f.t.
9.94	19.3
9.96	29.7
10.03	40.1
10.19	54.5

Rodebush , 1918

%	f.t.	%	f.t.
5.76	-3.48	18.20	-14.33
8.19	-5.17	18.69	-14.77
9.72	-6.32	19.84	-16.21
12.43	-8.52	21.69	-18.73
13.39	-9.41	22.90	-20.56
15.16	-11.04	-23.32	-21.12 E

Maass and Hatcher, 1922

%	f.t.	%	f.t.
0	-1.72	8.61	-7.67
1.08	-2.37	11.81	-10.57
3.14	-3.62	15.63	-10.37
5.42	-5.07	17.00	above 0

Findlay and Cruickshank, 1926

26.40 %   f.t. = 20°

Küpper, 1926

%	f.t.		%	f.t.
26.27	0	NaCl.2 aq.	26.86	50
26.29	5	NaCl.	26.96	55
26.32	10		27.01	57
26.36	15.3		27.05	59
26.38	17.9		27.07	60
26.39	20		27.09	61
26.40	20.6		27.16	64
26.45	25		27.30	70
26.47	27		27.33	71
26.51	30		27.55	80
26.53	31		27.63	83
26.60	35.5		27.81	90
26.67	40		28.15	100
26.84	49			

Scott and Frazier, 1927

26.40 %   f.t. = 25°

Neumann and Domke, 1928

molarity	f.t.
312.3	20
314.4	30
316.0	40

Chretien, 1929 (fig.)

%	f.t.	
0	0	Ice
9.1	-5	"
16.7	-14	
22.9	-21.6	NaCl.2aq + Ice
25.1	-3	
26.4	+0.2	NaCl.2aq + NaCl
28.3	108.8	b.t. satd.

Cornec and Neumeister, 1929

%	f.t.	%	f.t.
26.4	0	27.5	75
26.6	25	28.6	100
26.9	50		

Akhumov and Vasiliev, 1932

%	f.t.	%	f.t.
0	0	29.8	150
9.1	- 0.5	30.2	160
16.6	- 8	30.3	170
22.4	-21.2 E NaCl.2 aq	33.7	180
26.2	0	31.2	190
26.7	33.3	31.5	200
27.0	66.6	32.0	220
28.2	100	32.5	230
28.6	110	33.0	250
28.9	120	34.2	270
29.2	130	34.4	280
29.5	140	35.0	300

Cornec and Krombach, 1932

%	f.t.	%	f.t.
26.25	0	28.30	108.7
26.34	20	28.60	120
26.64	40	29.30	140
27.03	60	30.62	169.5
27.51	80	31.45	189.6
28.00	100		

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
36.1	285	45.6	388
37.3	301	47.5	409
40.6	334	49.1	430
42.6	356	51.3	455
43.3	364	100	800
45.1	384		

Khitrova, 1954

%	f.t.	
10	-6.2	ice
16	-11.6	"
19.6	-15.2	"
20.6	-16.4	"
22	-19.2	"
25	-8.6	NaCl.2aq.

Mun and Darer, 1956.

m	f.t.
2.46	- 8.80
1.80	- 6.25
1.11	- 3.85
0.64	- 2.19

Akhumov and Pylkova , 1956.

%	f.t.	limit t supercooling
28.20	100	26
28.88	121	48
29.78	149	73
30.85	179	106
32.01	208	132
32.39	218	145
33.17	236	163

Möller, 1862

%			
1 atm.	20 atm.	40 atm.	f.t.
26.25	26.35	26.44	0
26.32	26.38	-	9.0
26.35	-	-	12.0
26.30	26.39	26.40	15.0
26.35	26.37	-	20.0
26.37	26.47	-	25.0
26.47	26.53	-	30.0

von Stackelberg, 1896

%	P
18.5°	
26.4	1
27	500

Cohen, Inouye and Euwen, 1910 and Euwen, 1910

%	P	%	P
24.05°			
26.416	1	27.018	1000
26.605	250	27.198	1500
26.766	500	sat.sol.	

Sill, 1916

%	P Kg
25°	
26.44	1
26.58	250
26.72	500
26.82	750

Denecke, 1919

E	P kg	
-21.2	1	NaCl + Ice I
25.7	583	"
29.4	1037	"
33.9	1617	"
36.9	2110	"
38.7	2225	NaCl + IceI + IceIII
38.5	2248	NaCl + IceIII
35.5	2550	
35.0	2602	
66.0	2321	NaCl + IceI + IceIII
61.0	2249	
57.1	2217	
51.1	2186	
45.0	2205	
40.0	2187	
45.0	2199	
61.0	2198	

Adams and Hall, 1931

solubility increase (in%)	P kg
29.93°	
0.00	1
.26	296
.26	310
.32	416
.51	820
.53	886
.68	1270
.68	1278
.60	1230
.68	1306
.75	1451
.75	1443
.75	1433
1.01	1930

Keevil, 1942

mol%	t	P	mol%	t	P
12.0	183.0	7.27	20.5	385.7	145.0
12.4	205.1	11.76	21.7	410.0	178.9
13.0	230.2	19.29	23.8	442.5	230.1
13.5	246.7	25.08	25.1	467.5	269.0
13.8	254.6	27.51	27.0	485.5	294.4
15.5	299.3	56.03	28.5	514.2	335.1
16.9	327.3	78.50	33.5	550.5	370.1
17.8	344.4	96.29	41.1	600.0	388.7
18.3	354.3	106.4	50.5	646.2	368.5

Jänecke, 1949

%	tr.t.	P kg	%	tr.t.	P kg
26.3	0.05	1002	26.9	25.5	10020
27.3	11	2004	26.5	25	12024
27.6	18	4008	26.1	24	14028
27.5	22.5	6012	25.5	22	16032
27.3	25	8016			

$NaCl.2H_2O$ = NaCl + sat.sol.

Activity.

Olynyk and Gordon, 1943

m	activity NaCl	m	activity NaCl
	20°		
2.442	0.9156	4.028	0.8510
3.089	.8900	4.471	.8312
3.261	.8831	4.913	.8114
3.522	.8726	5.783	.7706
3.543	.8718	6.125[a]	.7546
	25°		
2.294	0.9210	3.973	0.8530
2.755	.9027	4.028	.8506
2.764	.9023	4.471	.8305
3.089	.8896	4.673	.8217
3.261	.8827	4.913	.8110
3.286	.8815	5.368	.7901
3.664	.8660	5.783	.7704
		6.145[a]	.7532
	30°		
2.442	0.9149	4.471	0.8296
3.089	.8892	4.659	.8217
3.261	.8822	4.913	.8108
3.286	.8810	4.966	.8082
4.028	.8500	5.783	.7701
		6.165 satd	.7522

Sheffer, Janis and Ferguson, 1939

m	activity	m	activity
	25°		
(6.145)	sat.sol.	0.8174	0.9729
6.142	0.7531	0.7066	0.9766
5.255	0.7936	0.5675	0.9813
4.229	0.8404	0.4158	0.9863
3.729	0.8628	0.3089	0.9898
2.770	0.9023	0.2031	0.9934
1.859	0.9367	0.1041	0.9966
1.488	0.9500	0.0508	0.99830
1.201	0.9598	0.02423	0.99918
0.9959	0.9669		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.932	1.6	0.962
0.2	0.925	1.8	0.972
0.3	0.922	2.0	0.983
0.4	0.920	2.5	1.013
0.5	0.921	3.0	1.045
0.6	0.923	3.5	1.080
0.7	0.926	4.0	1.116
0.8	0.929	4.5	1.153
0.9	0.932	5.0	1.192
1.0	0.936	5.5	1.231
1.2	0.943	6.0	1.271
1.4	0.951		

Properties of phases .

Density .

Karsten, 1846

%	d 0°	10°	15°	20°	40°	60°	80°
0	0.9999	0.9997	0.9992	0.9984	0.9923	0.9836	0.9737
0.5	1.0037	1.0034	1.0028	1.0019	0.9957	0.9869	0.9771
1	.0076	.0071	.0064	.0054	0.9991	0.9903	0.9805
1.5	.0114	.0108	.0100	.0090	1.0025	0.9937	0.9839
2	.0153	.0145	.0137	.0126	.0060	0.9971	0.9873
2.5	.0191	.0182	.0173	.0161	.0094	1.0005	0.9908
3	.0230	.0219	.0209	.0197	.0129	.0039	0.9942
3.5	.0268	.0256	.0246	.0233	.0163	.0073	0.9976
4	.0307	.0293	.0282	.0269	.0198	.0107	1.0010
4.5	.0345	.0330	.0319	.0305	.0233	.0141	.0044
5	.0384	.0367	.0355	.0341	.0267	.0175	.0079
5.5	.0422	.0405	.0392	.0378	.0302	.0210	.0113
6	.0461	.0442	.0429	.0414	.0337	.0244	.0148
6.5	.0500	.0479	.0466	.0450	.0373	.0279	.0182
7	.0538	.0517	.0503	.0487	.0408	.0314	.0217
7.5	.0577	.0554	.0540	.0523	.0443	.0348	.0251
8	.0616	.0592	.0577	.0560	.0478	.0383	.0286
8.5	.0654	.0629	.0614	.0597	.0514	.0418	.0320
9	.0693	.0667	.0651	.0633	.0550	.0453	.0355
9.5	.0732	.0705	.0688	.0670	.0586	.0589	.0390
10	.0771	.0742	.0726	.0707	.0621	.0524	.0425
10.5	.0810	.0780	.0763	.0744	.0658	.0559	.0460
11	.0849	.0818	.0801	.0781	.0694	.0595	.0495
11.5	.0888	.0856	.0838	.0819	.0730	.0631	.0530
12	.0927	.0894	.0876	.0856	.0766	.0666	.0565
12.5	.0966	.0933	.0914	.0894	.0803	.0702	.0601
13	.1005	.0971	.0952	.0931	.0840	.0738	.0636
13.5	.1044	.1009	.0990	.0969	.0877	.0775	.0671
14	.1083	.1048	.1028	.1007	.0914	.0811	.0707
14.5	.1123	.1087	.1066	.1045	.0951	.0848	.0743
15	.1162	.1125	.1105	.1083	.0988	.0884	.0778
15.5	.1202	.1164	.1143	.1122	.1026	.0921	.0814
16	.1241	.1203	.1182	.1160	.1063	.0958	.0850
16.5	.1281	.1242	.1221	.1199	.1101	.0995	.0886
17	.1321	.1281	.1260	.1237	.1139	.1032	.0923
17.5	.1361	.1321	.1299	.1276	.1177	.1070	.0959
18	.1401	.1360	.1338	.1315	.1216	.1107	.0995
18.5	.1441	.1400	.1378	.1355	.1254	.1145	.1032
19	.1481	.1439	.1417	.1394	.1293	.1183	.1069
19.5	.1521	.1479	.1457	.1433	.1332	.1221	.1105
20	.1562	.1519	.1497	.1473	.1371	.1259	.1142
20.5	.1602	.1559	.1537	.1513	.1410	.1298	.1179
21	.1643	.1600	.1577	.1553	.1449	.1336	.1217
21.5	.1684	.1640	.1617	.1593	.1489	.1375	.1254
22	.1724	.1681	.1657	.1633	.1529	.1414	.1292
22.5	.1765	.1721	.1698	.1674	.1569	.1453	.1329
23	.1806	.1762	.1739	.1714	.1609	.1492	.1367
23.5	.1848	.1803	.1780	.1755	.1649	.1532	.1405
24	.1889	.1845	.1821	.1796	.1690	.1572	.1443
25	.1972	.1927	.1904	.1879	.1772	.1652	.1520
26	.2056	.2011	.1987	.1963	.1855	.1733	.1597
26.4	.2089	.2045	.2021	.1996	.1888	.1765	.1628
26.8	-	-	.2055	.2030	.1922	.1798	.1660

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0127	16	1.1169
4	.0271	18	.1325
6	.0416	20	.1485
8	.0563	22	.1646
10	.0712	24	.1810
12	.0862	26	.1977
14	.1014	26.48	.2016

Grassi, 1851

%	t	d
15.323	18.5	1.1226
24.004	18.1	1.2024
24.004	39.6	1.188

Kremers, 1855 and 1857

t	d		
	4.5%	9.1%	13.3%
0	1.00362	1.00527	1.00632
10	.00217	.00285	.00327
19.5	.00000	.00000	.00000
30	0.99674	0.99629	0.99566
40	.99295	.99216	.99167
50	.98853	.98769	.98695
60	.98380	.98282	.98215
70	.97846	.97748	.97690
80	.97270	.97192	.97139
90	.96652	.96595	.96552
100	.95990	.95961	.95956
	17%	20.9%	23.9%
10	1.00355	1.00381	1.00397
19.5	1.00000	1.00000	1.00000
30	0.99566	0.99545	0.99528
40	.99129	.99098	.99080
50	.98665	.98628	.98607
60	.98173	.98137	.98116
70	.97656	.97626	.97604
80	.97114	.97094	.97081
90	.96547	.96542	.96533
100	.95956	.95959	.95978

%	d
19.514°	
6.402	1.0442
12.265	.0876
17.533	.1283
22.631	.1691
26.530	.2014

Buchholz, 1856

%	d	%	d
	18.75°		
3	1.0129	16	1.1168
4	.0273	18	.1324
6	.0418	20	.1483
8	.0564	22	.1643
10	.0712	24	.1807
12	.0861	26	.1973
14	.1014	26.63	.2026

Schiff, 1858, 1859 and 1860

%	d	%	d
	20°		
1	1.0048	15	1.1070
2	.0115	16	.1148
3	.0183	17	.1227
4	.0252	18	.1327
5	.0322	19	.1388
6	.0393	20	.1470
7	.0465	21	.1551
8	.0537	22	.1634
9	.0611	23	.1717
10	.0686	24	.1801
11	.0762	25	.1885
12	.0838	26	.1969
13	.0915	27	.2054
14	.0992		

Schmidt, 1859

%	d	%	d
	16°		
4.83286	1.035161	17.82833	1.134003
8.69344	.063805	21.37807	1.162295
14.80111	.110339	25.38901	1.196053

Fouqué, 1867

%	d		
	0°	20.6°	34.4°
0.1	1.0029	1.0016(21.6°)	-
0.9	.0073	.0052	1.0013
1.9	-	.0124	.0083
4.0	1.0353	.0324	.0277
9.1	.0716	.0676	.0626
17.4	.1343	.1275(18.4°)	.1206(34.2°)
25.3	.2029	.1928(18.4°)	.1850(34.6°)

van der Willigen, 1869

%	d	%	d
	25.75°		
8.65	1.05794	21.69	1.15785
15.85	.11194	22.78	.16731
16.61	.11745	26.58	.19845
20.73	.15019		

Lundquist, 1869

%	t	d
0	40.3	0.9921
23.43	43.9	1.1670

Thomsen, 1871

mol%	d	mol%	d
	18°		
9.1	1.1872	2.0	1.0444
4.8	.1033	1.0	1.0234
3.2	.0718	0.5	1.0118

Marignac, 1871

t	d	t	d	t	d
7.5 mol%		3.8 mol%		2 mol%	
0	1.16188	0	1.08867	0	1.04688
8.87	.15802	8.72	.08603	8.20	.04537
14.28	.15660	16.15	.08353	14.41	.04387
19.65	.15309	15.22	.08382	19.25	.04248
15.27	.15512	22.57	.08106	25.45	.04047
19.34	.155323	27.62	.07905	28.35	.03944
25.58	.15024	28.84	.07853		
29.34	.14841				
1 mol%		0.5 mol%			
0	1.02389	0	1.01203		
4.00	.02360	4	.01193		
10.75	.02268	8.60	.01157		
13.80	.02210	10.75	.01127		
18.66	.02101	18.66	.00985		
26.20	.01891	26.20	.00792		

Kohlrausch and Grotrian, 1875

%	d	%	d
	18°		
5	1.0346	20	1.1482
10	.0710	24	.1802
15	.1089		

Guthrie, 1875

%	d	%	d
	26°		
3.94	1.0246	13.13	1.0725
5.25	.0284	14.45	.0800
6.57	.0357	15.76	.0874
7.88	.0428	17.07	.0963
9.19	.0501	18.39	.1087
10.51	.0574	26.27	.2011
11.82	.0650		

Kohlrausch, 1879

%	d	%	d
	18°		
5	1.0345	25	1.1898
10	.0707	26	.1982
15	.1087	26.4	.2014
20	.1477		

Rönnberg, 1880

t	d				
	5%	10%	15%	20%	25%
4	1.03510	1.06735	1.09760	1.12616	1.15361
5	.03495	.06715	.09735	.12585	.15325
10	.03420	.06613	.09611	.12431	.15143
15	.03315	.06483	.09454	.12248	.14933
20	.03188	.06332	.09274	.12037	.14697
25	.03035	.06150	.09061	.11791	.14423

Volkmann, 1882

%	d	%	d
20°		15.16°	
25.48	1.1932	25.86	1.1987
21.44	.1596	15.62	.1154
14.39	.1074	6.46	.0463
10.11	.0720		
5.25	.0360		

Bender, 1883

M	d	M	d
	15°		
1	1.0401	4	1.1522
2	1.0788	5	1.1888
3	1.1164		

Nicol, 1884

mol%	d	
	20°	40°
0.5	1.00966	-
1.0	.02074	-
1.95	.04205	1.01397
2.9	.06248	-
3.85	.08216	-
4.75	.10081	1.09178
5.65	.11900	-
6.55	.13636	-
7.40	.15298	-
8.25	.16932	-
9.10	.18490	-
9.90	.19978	1.04453

Rother, 1884

%	d	%	d
	15°		
5.06	1.0358	20.01	1.1511
10.01	1.0723	25.24	1.1941
14.94	1.1104		

Nicol, 1884

t	d
10.5 N (at 20°)	
20	1.19774
40	.17984
98	.10519

Fink, 1885

%	d	%	d
	18°		
0.99	1.0059	19.60	1.1446
4.99	.0350	24.90	.1885
9.90	.0694	26.39	.2016
14.82	.1078		

Traube, 1885

%	d
15°	
9.34	1.0703
17.65	1.1331

Schumann, 1887

%	d	%	d
	15°		
0	0.9991	18.18	1.1356
1.32	1.0086	22.16	1.1678
3.51	1.0245	26.21	1.2017
13.53	1.0992		

Tait, 1888

%	d		
	0°	6°	12°
17.6358	1.138467	1.136040	1.133565
13.3610	.101300	.099341	.097244
8.3078	.067589	.066144	.064485
3.8845	.029664	.028979	.027935

Leblanc, 1889

%	d
20°	
0	0.99823
4.79	1.03234
5.31	1.03637
24.13	1.18200

Wegner, 1889

%	d	%	d
	20°		
7.9879	1.055840	22.4036	1.167226
16.5657	.120589	23.6326	.177286
21.3677	.158737	24.7345	.186480

Buchkremer, 1890

%	d	%	d
	20°		
0	0.99827	4.0140	1.02941
1.396	1.01051	13.148	.09814
2.2625	.01638	23.800	.188440
3.0118	.02250		

Schütt, 1890

%	d	%	d
	18°		
25	1.18911	2	1.01291
20	.14821	1	.00579
15	.10914	0.5	.00220
10	.07124	0.3	.00079
5	.03454	0.1	.99936
3	.02007	0	.99866

Bodländer, 1891

%	d	%	d
13°		15.5°	
26.26	1.2020	26.43	1.2015
26.36	1.2025		

Brückner, 1891

mol%	d	mol%	d
	15°		
0	0.9992	3	1.1160
0.5	1.0196	3.5	1.1343
1	1.0398	4	1.1524
1.5	1.0592	4.5	1.1702
2	1.0792	5	1.1877
2.5	1.0974		

Charpy, 1893

%	d
0°	
23.0821	1.1820
19.1932	1.1501
14.3415	1.1110
9.4120	1.0721
5.1536	1.0393

Perkin, 1894

%	d	
	10°	20°
26.174	1.2025	1.1974

Delaite, 1894

%	d	%	d
	15°		
20.25	1.1519	0.36	1.0014
10.81	.0791	.18	1.0000
5.61	.0400	.09	0.9994
2.86	.0157	.05	0.9992
1.45	.0093	0.00	0.9991
0.73	.0041		

Gilbault, 1897

%	d	%	d
	20°		
0	0.9982	14	1.1019
2	1.0127	17	.1250
5	.0344	20	.1491
8	.0566	23	.1737
11	.0791	26.22	.2008

Sentis, 1897

mol%	t	d
8	11.75	1.1665
8	16.3	.1704
7	15.25	.1456
6	15.4	.1267
5	10.0	.1088
4	13.7	.0869
3	17.0	.0654
2	17.25	.0436
1	16.3	.0215
1	19.4	.0210

Barnes and Scott, 1898

%	d	%	d
	18°		
25.37	1.1928	4.810	1.0332
21.25	.1592	2.991	.0202
17.35	.1277	2.593	.0173
13.25	.0958	1.746	.0111
9.34	.0665	0	.9986

Linebarger, 1899

%	d	%	d
25°		20°	
5.62	1.0370	5.62	1.0388
10.87	.0726	10.87	.0746
14.52	.0944	15.72	.1146
15.72	.1123	26.19	.20046

Lyle and Hosking, 1902

t	d			
	0	0.1 N	0.2 N	0.5 N
	(molarity at 15°)			
0	0.9999	1.0048	1.0095	1.0220
10	.9997	.0043	.0087	.0210
15	.9992	.0036	.0080	.0201
18	.9986	.0032	.0075	.0195
20	.9983	.0028	.0071	.0190
30	.9958	.0002	.0044	.0160
40	.9923	0.9966	.0005	.0123
50	.9882	.9924	0.9962	.0079
60	.9834	.9876	.9915	.0029
70	.9779	.9821	.9862	0.9975
80	.9719	.9763	.9805	.9916
90	.9656	.9700	.9742	.9854
100	.9586	.9632	.9675	.9792
	1.0 N	2.0 N	4.0 N	
0	1.0442	1.0841	1.1600	
10	.0421	.0809	.1554	
15	.0408	.0792	.1530	
18	.0401	.0782	.1517	
20	.0394	.0774	.1507	
30	.0361	.0734	.1457	
40	.0320	.0690	.1405	
50	.0275	.0643	.1351	
60	.0224	.0588	.1295	
70	.0170	.0533	.1236	
80	.0111	.0474	.1176	
90	.0049	.0414	.1113	
100	0.9983	.0353	.1049	

Bender, 1902

%	d	%	d
15°			
2.87	1.0200	15.72	1.1164
5.62	.0401	20.31	.1522
10.85	.0788	22.50	.1698

Berkeley, 1904

%	t	d	%	t	d
26.34	0.35	1.2090	27.16	61.70	1.1823
26.40	15.20	.2020	27.44	75.65	.1764
26.55	30.05	.1956	27.81	90.50	.1701
26.79	45.40	.1891	28.39	107.0 (b.t.)	.1631

Grabowsky, 1904

%	d		
	10°	18°	30°
0	0.999731	0.99863	0.99567
10.24	1.0824	1.0728	1.0679
18.71	1.1420	1.1383	1.1327
26.24	1.2067	1.2018	1.1945

Dinkhauser, 1905

%	d	%	d
18°			
0	0.9986	10.856	1.0778
2.870	1.0192	15.729	1.1157
4.262	1.0293	20.313	1.1520
5.629	1.0392	24.644	1.1871

Agerer, 1905

%	d
18°	
4.889	1 0351
22.639	1.1714

Demolis, 1906

t	d 295.5 g/l	227.8 g/l	149.6 g/l	176.1 g/l
0	1.1956	1.1545	1.1027	1.0546
10	.1912	.1510	.1006	.0528
20	.1864	.1464	.0978	.0500
30	.1812	.1414	.0931	.0465
40	.1755	.1360	.0880	.0420
50	.1695	.1320	.9823	.0370
60	.1632	.1245	.0765	.0318
70	.1571	.1185	.0706	.0261
80	.1511	.1128	.0646	.0205

g/l	d 0°	30°	50°	80°
0	0.9999	0.9957	0.9881	0.9718
50	1.0382	1.0299	1.0206	1.0050
75	.0536	.0458	.0362	.0198
100	.0699	.0617	.0518	.0396
125	.0862	.0776	.0671	.0499
150	.1029	.0936	.0827	.0650
175	.1199	.1092	.0982	.0804
200	.1364	.1247	.1134	.0959
225	.1526	.1397	.1287	.1111
250	.1682	.1546	.1433	.1256
275	.1834	.1691	.1576	.1397
300	.1982	.1840	.1720	.1537

Getmann, 1907

%	d
18°	
9.38	1.0651
13.46	.1454
25.07	.1963

Chéneveau, 1907

%	d	%	d
15°			
25.65	1.1976	14.07	1.1041
23.46	.1794	11.36	.0835
21.21	.1611	8.70	.0638
18.90	.1425	5.92	.0429
16.46	.1232	2.99	.0217

Guerdjikova, 1910

%	d	%	d
25°			
0	-	18.300	1.1318
11.410	1.0882	24.292	.1794

Baxter, Boylston and al., 1911

%	d	%	d
25°			
0.5280	1.00079	1.1068	1.00477
.5493	.00098	5.3562	.03488
.9980	.00413	5.4131	.03532
1.0618	.00432	14.344	.10146

Clausen, 1912

M	d	M	d
0°			
0.884	-	3.924	1.1562
1.330	1.0771	5.085	1.1972
2.843	1.1162	5.325	1.2056

Grufki, 1913 and Lübben, 1913

M	%	d
18°		
0	0	0.99862
0.50	2.89	1.01951
1.012	5.67	1.03967
2.009	10.89	1.07817
4.003	20.32	1.15145

Herz, 1914

M	d
25°	
5.0	1.1831
3.75	.1382
2.50	.0936
1.25	.0454

Walker, 1914

%	d	%	d
15.5°			
0	0.999	18	1.133
6	1.041	24	.164
12	1.087	28	.197

Baxter and Wallace, 1916

%	d	%	d	%	d
0°		25°		50.04°	
25.08	1.19930	25.29	1.18851	25.52	1.17656
16.57	.10483	13.67	.09625	13.81	.08527
5.75	.04397	5.78	.03807	5.83	.02804
2.86	.02230	2.94	.01771	2.96	.00812
1.47	.01117	1.48	.00747	1.49	.99818

Herz, 1917

M	d
25°	
5.000	1.1827
3.750	.1380
2.500	.0940
1.250	.0452
0	.9971

R.Bousfield and E.Bousfield, 1919

%	d	%	d
18°			
0	0.99862	14	1.10171
1	1.00578	15	.10939
2	.01296	16	.11713
3	.02015	17	.12494
4	.02736	18	.13282
5	.03460	19	.14076
6	.04188	20	.14876
7	.04920	21	.15682
8-	.05656	22	.16494
9	.06397	23	.17313
10	.07142	24	.18140
11	.07893	25	.18975
12	.08648	26	.19818
13	.09407		

Bousfield, 1919

%	d 7°	d 20°	d 33°
25.000	1.19538	1.18879	1.181885
18.945	.14531	.13952	.13323
13.519	.10238	.09745	.09178
10.005	.07519	.07094	.06576
5.625	.04183	.03857	.03404
2.8669	.02129	.01873	.01468
1.4476	.010735	.00858	.00482
0	.99993	.99823	.99473

Pulvermacher, 1920

M	d	M	d
25°			
5.149	1.1878	2.574	1.0961
4.505	.1653	1.278	.0481
3.862	.1428	0.257	.0075

Wasastjerna, 1920

M	d	M	d
	18°		25°
0.0000	0.99862	0.0000	0.99707
.1000	1.00278	.0998	1.00123
.2000	.00686	.1997	.00523
.5000	.01915	.4990	.01729
.7000	.02726	.6986	.02528
1.0000	.03914	.9979	.03699
.5000	.05870-	1.4966	.05633
2.0000	.07776	.9950	.07509
.5000	.09647	2.4936	.09365
3.0000	.11468	.9922	.11179
.5000	.13321	3.14903	.13008
4.0000	.15083	-	-
.5000	.16899	4.4871	1.16565
5.0000	.18632	.9851	.18278

Hall, 1924

%	d	%	d
25°			
0.99673	1.004106	14.8226	1.105057
2.0210	.011325	19.9310	.144889
4.8249	.031241	23.9395	.177362
4.8176	.031198	24.9320	.185587
9.9955	.068807	24.9205	.185502
9.9849	.068732	25.5098	.190405
14.8220	.105050		

Koch, 1924

d	t	d	t	d	t
7.634%		18.88%		22.15%	
1.0586	-2.3	1.1525	-14.2	1.1813	-16.0
.0578	+2.2	.1522	13.6	.1809	14.9
.0568	6.7	.1494	-5.9	.1802	13.3
.0551	13.1	.1460	+2.4	.1783	8.9
.0549	13.9	.1424	11.1	.1767	-5.1
.0541	16.7	.1398	16.9	.1734	+2.2
.0536	18.3	.1376	21.4	.1701	9.2
.0530	20.0			.1660	17.7
		26.17%		.1648	20.1
14.00%		1.2083	-0.4	.1646	20.5
		.2077	+0.7		
1.1106	-8.4	.2064	3.3		
.1099	-6.2	.2038	8.4		
.1080	+0.3	.2018	12.4		
.1049	8.8	.1992	17.2		
.1023	15.7	.1976	20.2		
.1009	19.5				

Manchot, Jahstorfer and Zepter, 1924

%	d	%	d
	25°		
6.723	1.0438	6.548	1.0420
13.504	.0874	13.212	.0850
25.255	.1600	26.424	.1660

Herz and Martin, 1924

t	d	t	d
	26.694 g/100cc		
20	1.1825	60	1.1619
30	1.1777	70	1.1562
40	1.1726	80	1.1503
50	1.1672	90	1.1441

Goard, 1925

M	d
20°	
1	1.036
2	.077
3.22	.110
5	.118

Rakshit, 1925

%	d
20°	
1	1.00572
5	.03296
10	.06613
30	.19009

De Block, 1925

%	d	%	d
	16°		
0	0.9990	16.5	1.1222
5.62	1.0403	25.2	1.1971
10.8	1.0791		

Gerlach, 1926

%	d	%	d
	15°		
0	0.99913	14	1.10288
1	1.00638	15	.11049
2	.01362	16	.11840
3	.02085	17	.12632
4	.02809	18	.13424
5	.03534	19	.14215
6	.04275	20	.15006
7	.05016	21	.15830
8	.05759	22	.16653
9	.06500	23	.17477
10	.07241	24	.18300
11	.08003	25	.19124
12	.08764	26	.19993
13	.09526		

Geffcken, 1929

M	d	M	d
	25°		
0	0.99707	3.0568	1.10752
1.9088	1.06907	3.9399	.13658
1.9886	.07185	4.9515	.16488
1.9940	.07204	4.9962	.16624
2.0175	.07287	6.0551	.19550
2.9992	.10567		

Herz and Hiebenthal, 1929

M	d 25°	d 70°
0.5	1.0175	0.9986
1.0	.0369	1.0176
2.0	.0746	.0176
3.0	.1084	.0866
4.0	.1472	.1233

Scott, Obenhaus and Wilson, 1934

%	d	%	d
	35°		
24.8200	1.17942	6.1298	1.13671
16.9863	.11690	2.8661	.11388
9.1598	.15852		

Spacu and Popper, 1934

%	d	%	d
	20°		
25.838	1.19588	11.682	1.08310
24.850	.18780	0	.99823
17.571	.12844		

Usanovich and Mun, 1955

%	d	%	d
	25°		
5.77	1.0375	19.02	1.1373
10.39	.0710	25.00	.1865
14.57	.1026		

Herz, 1930

%	d	%	d
	18°		
2.870	1.0192	15.729	1.1157
4.262	1.0293	20.313	1.1520
10.856	1.0778	24.644	1.1871

Hölemann and Kohner, 1931

M	d	M	d
	25°		
0	0.99707	3.4476	1.11993
1.6692	1.06062	3.4477	1.11998
1.6693	1.06063	5.4140	1.17813
1.6693	1.06065		

M	d	M	d
35.0°		45.0°	
0	0.99406	0	0.99024
1.6689	1.05656	1.6689	1.05207
1.6690	1.05660	1.6690	1.05207
3.4494	1.11538	3.4497	1.11030
5.4144	1.17282	3.4498	1.11029
		5.4186	1.16743

Rodnyanski and Galinker, 1955 (fig.)

t	d 1 N	d 2 N	d 3 N
25	1.036	1.072	1.101
50	1.024	1.059	1.088
100	0.992	1.028	1.057
150	0.954	0.990	1.022
200	0.910	0.951	0.986
250	0.862	0.906	0.947
300	0.804	0.858	0.902
340	0.734	0.795	0.838

Gellings, 1956

N	d	N	d
	25°		
.18611	1.00766	1.8886	1.09896
.46425	.02282	2.4813	.12965
.69031	.03522	3.0487	-
.95575	.04971	.4976	1.18097
1.1612	.06076	4.0068	.20669
.6157	.08463	.4912	.23060
		.8497	.24831

Density sat. sol.

Andreae, 1884

t	d (sat.sol.)	t	d
15	1.20251	50	1.18753
20	1.20034	60	1.18333
30	1.19602	70	1.17918
40	1.19176	80	1.17506

Chernaya, 1912

%	t	d	%	t	d
27.28	70	1.178	26.37	20	1.202
27.04	60	.184	26.30	10	.205
26.83	50	.189	26.27	0	.210
26.64	40	.193	24.75	-10	.200
26.49	30	.198			

Scott and Frazier, 1927

26,40 % d = 1,19796

Hrynakowski, 1927

sat. sol. 49.80° d = 1,1925

Flöttmann, 1928

%	t	d
26.343	15	1.2024
26.403	20	1.2001
26.47	25	1.1979

Cornec and Neumeister, 1929

%	t	d
26.4	0	1.209
26.6	25	1.198
26.9	50	1.185
27.5	75	1.176
28.6	100	1.164

Cornec and Krombach, 1932

%	t	d
26.25	0	1.209
26.34	20	1.201
26.64	40	1.192
27.03	60	1.184
27.51	80	1.175
28.00	100	1.166
28.30	108.7	1.162

Expansion coefficients

Sorby, 1859

25%

$$vol = 1.0000 + 0.000397\ t + 0.00000127\ t^2$$

DeHeen, 1881

t	vol.	t	vol.	t	vol.
24.51%		16.92%		9.74%	
10.00	1.000000	10.00	1.000000	10.00	1.000000
13.30	.001320	15.20	.001965	16.26	.001821
18.30	.003480	20.77	.004133	20.61	.003247
23.56	.005789	25.90	.006046	25.48	.004840
30.50	.008854	30.52	.007945	32.20	.007473
36.58	.011638	36.60	.010692	40.30	.010917
42.72	.014597	42.67	.013412	50.58	.015876
51.16	.018866	50.29	.017133	57.25	.019464
59.10	.022983	58.60	.021179		
67.43	.027459	65.55	.024950		
75.20	.031996	71.20	.027958		
		75.62	.030553		

Nicol, 1884, 1886 and 1887

mol%	molar vol.sol.
20°	
1.95	1836.75
3.85	1876.93
5.65	1920.06
7.4	1963.83
9.1	2010.70

t	vol.	t	vol.
1.95 mol%		3.85 mol%	
20	100	20	100.000
45.4	100.954	45.7	101.059
50.4	101.188	50.9	101.300
56.5	101.495	56.5	101.582
61.7	101.758	61.5	101.839
67.6	102.079	67.6	102.168
72.5	102.364	72.5	102.445
78.0	102.699	78.8	102.812
5.65 mol%		7.4 mol%	
20	100.000	20	100.000
45.1	101.096	45.1	101.132
50.5	101.355	50.5	101.398
56.2	101.633	56.5	101.690
61.4	101.901	61.6	101.953
67.2	102.214	67.5	102.271
72.5	102.495	72.0	102.518
78.1	102.826	78.0	102.848
9.1 mol%		9.1 mol%	
20	100.000	61.7	101.989
45.6	101.178	67.1	102.279
50.9	101.440	72.3	102.555
56.3	101.715	78.2	102.878

Compressibility .

Grassi, 1851

%	t	π
15.323	18.5	321
24.004	18.1	257
24.004	39.6	263

Schumann, 1887

%	π	%	π
	15°		
0	46.4	18.18	33.3
1.32	45.3	22.16	28.2
3.51	44.0	26.21	25.5
13.53	33.9		

Tait, 1888

%	π		
	157 kg	314 kg	471 kg
		0°	
0	50.3	49.0	47.7
3.9	44.9	43.8	42.8
8.8	39.6	38.6	37.8
13.4	35.4	34.5	33.8
17.63	32.1	31.3	30.6

Gilbault, 1897

%	π (1-300 atm.)	%	π (1-300 atm.)
	20°		
0	44.37	14	33.48
2	42.57	17	31.65
5	40.03	20	29.91
8	37.67	23	28.28
11	35.52	26.22	26.69

Pohl, 1906

%	π (H_2O at 13.9° = 1)
13.9°	
5.04	0.885
8.61	.799
12.60	.736
14.82	.729
15.51	.713
17.50	.687
19.86	.625
25.18	.583
satd.	.542

Adams, 1931

K = Dv/V_0 at P

P bars	K 0%	K 5%	K 10%
		25°	
1	-0.0000	-0.0000	-0.0000
500	.0212	.0192	.0776
1000	.0393	.0361	.0331
1500	.0555	.0513	.0472
2000	.0699	.0647	.0597
3000	.0945	.0880	.0817
4000	.1152	.1076	.1000
5000	.1330	.1248	.1166
6000	.1485	.1397	.1312
7000	.1622	.1530	.1442
8000	.1746	.1647	.1555
9000	.1858	.1759	.1662
10000	.1964	.1860	.1762
11000	1.2059	.1953	.1853
12000	1.2147		
	15%	20%	25%
1	-0.0000	-0.0000	-0.0000
500	.0161	.0147	.0133
1000	.0304	.0277	.0251
1500	.0433	.0396	.0361
2000	.0551	.0507	.0462
3000	.0756	.0699	.0644
4000	.0933	.0864	.0799
5000	.1089	.1015	.0941
6000	.1228	.1147	1.1070
7000	.1353	.1267	.1183
8000	.1462	.1375	.1289
9000	.1567	.1479	.1388
10000	.1668	.1574	.1481
11000	.1757	.1661	.1564
12000	-	.1740	.1643

Freyer, 1931

%	π	%	π
	20°		
1	45.00	16	32.45
6	40.28	20	29.76
10	36.96	24	27.36

t	π	t	π
	26%		
20	26.25	30	26.22
25	26.23	40	26.32

Scott, Obenhaus and Wilson, 1934

%	π	%	π
		35°	
2.8661	39.71	16.9863	30.99
6.1298	37.54	24.8200	27.10
9.1598	35.56		

Dörsing, 1908

%	velocity of sound (m/s)
	15°
0	1249
10	1251.2
15	1280.6
20	1371.7

Busz, 1938

%	velocity of sound (m/s)
	14-16°
0	1480
5	1530
10	1590
13.2	1630
20	1720
26.4	1790

Kurochkin, 1929

g/100cc H_2O	Dv
4	1.12
10	3.20
20	6.00

Dv = % volume increase by dissolution.

Viscosity .

Grotrian, 1876

%	t	η	t	η	t	η
4.95	8.32	1989	14.74	1641	22.03	1316
15.00	7.94	2569	14.41	2092	21.99	1702
23.86	7.87	3643	14.55	2951	21.30	2320

%	η	$\tau.10^5$
	18°	
4.97	1496	309
15.00	1907	276
23.86	2629	359

Brückner, 1891

mol%	η 15°	η 20°
0	1143.9	1008.6
0.5	1182.6	1048.2
1	1235.5	1096.2
1.5	1295.1	1151.3
2	1365.1	1214.7
2.5	1447.5	1287.8
3	1543.5	1374.5
3.5	1651.8	1468.2
4	1783.3	1583.9
4.5	1931.5	1708.9
5	2117.8	1867.3

Lyle and Hosking, 1902

t	η 0	η 0.1 N	η 0.2 N
	(molarity at 15°)		
0	1794	1805	1805
10	1309	1312	1312
15	1143	1147	1147
18	1060	1063	1063
20	1009	1013	1013
30	802	805	805
40	657	657	685
50	553	553	553
60	472	472	472
70	407	407	407
80	358	358	358
90	316	316	316
100	285	285	285

t	η 0.5 N	η 1.0 N	η 2.0 N	η 4.0 N
	(molarity at 15°)			
0	1820	1851	2039	2676
10	1342	1392	1506	2014
15	1177	1223	1328	1762
18	1092	1135	1237	1636
20	1041	1082	1182	1557
30	837	869	956	1254
40	685	715	788	1034
50	579	599	668	873
60	493	515	578	746
70	427	447	502	647
80	375	394	442	569
90	333	349	397	506
100	300	314	355	456

Getmann, 1907

%	η
	18°
9.38	1259
18.46	1559
25.07	2030

Hosking, 1909

t	η 0%	1%	5%	10%	20%
0	1794	1799	1862	2041	2666
5	1520	1534	1600	1738	2269
10	1309	1320	1390	1528	1968
15	1143	1154	1223	1345	1733
20	1009	1018	1083	1194	1540
25	897	905	973	1066	1378
30	802	810	876	961	1239
35	724	731	796	875	1124
40	657	666	725	802	1025
45	601	609	666	737	939
50	553	559	614	680	866
55	510	519	569	633	801
60	472	481	531	589	744
65	437	448	492	550	693
70	407	416	456	516	649
75	382	391	425	484	607
80	360	368	399	455	571
85	339	346	375	427	537
90	320	325	355	401	507

Walker, 1914

d	η (water=100)	d	η (water=100)
	0°		
1.19	187.0	1.10	122.0
1.15	150.0	1.05	107.0
	10°		
1.18	128.5	1.10	91.5
1.15	111.0	1.05	79.1
	30°		
1.172	74.5	1.10	57.0
1.15	68.0	1.05	50.0
	50°		
1.16	47.8	1.10	37.9
1.15	45.0	1.05	34.0
	100°		
1.15	21.5	1.05	16.6
1.10	18.7		

d	η	d	η
	17°		
0.999	1085	1.064	1235
1.007	1101	.077	1283
.014	1119	.094	1336
.024	1140	.114	1453
.036	1171	.141	1556
.057	1220	.169	1750

Herz, 1914 and 1917

molarity	η
	25°
5.000	1666
3.750	1371
2.500	1156
1.250	1000
0	895

Pulvermacher, 1920

M	η (water = 1)
	25°
5.149	1.895
4.505	1.700
3.862	1.538
2.574	1.303
1.278	1.126
0.257	1.021

Giordani, 1924

M	η 14°	34.3°	45.5°
1.093	1215.9	762.8	646.7
2.186	1382.6	-	722.2
3.279	1589.3	1051.3	851.2
4.372	1869.8 (13°6)	1213.3	974.3
5.465	2256.9 (13°7)	1409.5	1119.9

Erk, 1924

t	η 10.017 %	19.95 %
25	1071	1388
20	1191	1560
18	1248	1633
16	1307	1714
14	1369	1801
12	1441	1894
10	1517	1997
8	1600	2107
6	1794	2229
4	1904	2362
2	2019	2510
0	2153	2675
- 2	2294	2850
- 4	2449	3040
- 6	-	3151
- 6.7	2509	-
- 8	-	3489
-10	-	3774

Herz and Martin, 1924

t	η	t	η
	26.694 g/100 cc		
20	1835	60	877
30	1471	70	767
45	1210	80	682
50	1020	90	635

Hrynakowski, 1927

sat. sol. 49.80° η = 1153

Stakelbeck and Plank, 1929

%	η 20°	10°	0°	- 5°	-10°	-15°	-20°
0.9	1025	1335	1825	-	-	-	-
1.9	1033	1345	1830	-	-	-	-
2.9	1043	1355	1855	-	-	-	-
3.9	1055	1375	1845	-	-	-	-
4.9	1069	1395	1860	-	-	-	-
5.9	1083	1415	1885	-	-	-	-
7.0	1099	1440	1915	-	-	-	-
8.0	1119	1465	1950	2342	-	-	-
9.1	1139	1495	1990	2395	-	-	-
10.2	1159	1525	2030	2450	-	-	-
11.3	1184	1560	2080	2512	-	-	-
12.3	1214	1595	2132	2575	-	-	-
13.4	1249	1638	2195	2650	3275	-	-
14.5	1284	1681	2218	2735	3385	-	-
15.6	1326	1734	2353	2822	3505	-	-
16.7	1368	1789	2438	2930	3645	-	-
17.9	1419	1849	2539	3070	3815	4720	-
19.0	1465	1909	2643	3210	3985	4925	-
20.1	1520	1979	2755	3360	4180	5155	-
21.2	1580	2054	2875	3522	4395	5395	-
22.3	1653	2139	3005	3695	4610	5665	6975
22.4	1655	2144	3025	3709	4630	5690	-
23.5	1732	2242	3152	3892	4885	5980	-
24.7	1822	2362	3310	4105	5205	-	-
25.9	1932	2511	3500	-	-	-	-
27	1970	2557	3590	-	-	-	-

Giordani and Maresca, 1929

t	η	t	η	
N/1		150g/1		
17.6	1164	16.4	1415	all
35.8	789	34.8	973	concen-
44.9	671	45.8	819	trations
65.4	492	65.2	607	at 20°
250g/1		300g/1		
20.2	1620	20.7	1796	
35.1	1193	35.4	1314	
44.8	1004	44.8	1102	
65.0	757	65.4	825	

Herz and Hiebenthal, 1929

M	η 25°	70°
0.5	938.1	434.4
1.0	978.0	459.4
2.0	1082.4	514.5
3.0	1210.5	574.5
4.0	1395.5	649.1

Surface tension .

Volkmann, 1882

%	σ	%	σ
20°		15-16°	
25.48	82.6	25.86	83.2
21.44	80.6	15.62	78.7
14.39	77.5	6.46	75.6
10.11	76.1		
5.25	74.8		

Rother, 1884

%	σ	%	σ
	15°		
5.06	73.43	20.01	78.06
10.01	74.71	25.24	80.37
14.94	76.35		

Traube, 1885

%	σ	%	σ
	15°		
9.34	74.91	17.65	78.29

Ochse, 1890

t	σ 0%	t	σ 5 g/100 cc aq.	t	σ 10 g/100 cc aq.
0	80.65	3.5	68.69	4	60.17
4	77.51	-	-	-	-
8	75.36	10	65.75	-	-
15	72.91	15.5	63.11	15	57.23
25	70.16	25	61.05	25	55.66
40	64.19	40	57.42	35	54.19
55	60.27	55	53.80	45	53.11
				55	51.64

t	σ 15g /100 cc aq.	t	σ 20g /100 cc aq.
3.5	55.46	4	51.05
15	52.71	-	-
26	52.62	15	50.66
35	52.03	25	50.37
45	51.54	35	49.78
55	51.05	45	49.39
		55	48.80

Sentis, 1897

mol%	t	σ	mol%	t	σ
8	11.75	82.4	4	13.7	77.6
8	16.3	81.9	3	17.0	76.2
7	15.25	80.5	2	17.25	75.3
6	15.4	79.4	1	16.3	74.4
5	10.0	79.8	1	19.4	74.2

Linebarger, 1899

%	σ	%	σ
25°		20°	
5.62	71.59	5.62	72.38
10.87	73.04	10.87	73.78
13.52	73.71	15.72	75.87
15.72	75.04	26.19	80.60

Forch, 1899

M	t	σ
	18°	
4.008	15.87	79.33
4.008	21.73	-
3.200	19.11	78.17
1.999	15.43	76.12
1.999	23.73	-
1.600	18.60	75.47
1.1972	17.80	74.85
0.7984	18.37	74.14
0.3987	14.55	73.65
0.3987	21.52	-

Grabowsky, 1904

%	σ	
	10°	30°
0	74.0	71.02
10.24	76.98	74.10
18.71	80.12	77.44
26.24	83.83	81.01

Forch, 1905

M	t	σ
4.006	14.0	84.29
3.204	15.0	82.71
2.085	14.2	80.83
1.600	14.9	79.86
0.484	14.6	78.19

Morgan and Bole, 1913

M	t	σ
0.5	0.0	76.68
"	34.3	71.27
1	0.0	77.34
"	34.2	72.09
2	0.0	78.81
"	33.2	73.69
2.5	0.0	79.64
"	33.9	74.65

De Block, 1925

%	σ	%	σ
	16°		
0	73.11	16.5	78.23
5.62	74.66	25.2	82.71
10.8	76.15		

Goard, 1925

M	σ
20°	
1	74.78
2	76.51
3.22	78.73
5	82.20

Kleinmann, 1926

Variation of σ with time (see author)

Hrynakowski, 1927

sat. sol. σ = 94.25 49.8°

Kremers, 1857

%	t	n_D
15.03	18	1.3581
26.15	17	1.3786

Börner, 1869

t	$n_{H\alpha}$	t	$n_{H\beta}$	t	$n_{H\gamma}$
		9.9933%			
44.3	1.343387	45.2	1.349682	44.4	1.353364
40.85	.343980	40.4	.350505	39.25	.354221
35.0	.344914	35.4	.351328	33.9	.355114
30.35	.345560	30.75	.352042	29.35	.355827
25.15	.346421	25.6	.352829	24.8	.356504
		19.9867%			
44.4	1.356397	43.5	1.363511	45.2	1.366985
39.8	.357217	39.1	.364256	40.5	.367834
34.6	.358036	34.15	.365107	34.0	.368967
28.9	.358997	29.65	.365904	29.65	.369674
24.7	.359709	24.8	.366666	24.9	.370522
		29.9805%			
44.5	1.367834	44.05	1.375301	45.8	1.379012
39.15	.368790	38.7	.376269	40.15	.379996
35.35	.369426	34.8	.376938	34.75	.380996
30.7	.370274	30.2	.377747	29.9	.381961
24.9	.371263	24.7	.378731	25.05	.382837

van der Willigen, 1869

s.l.	n			
	8.65%	15.85%	16.61%	20.73%
		25.75°		
A	1.34264	1.35520	1.35670	1.36398
a	.34355	.35617	.35756	.36493
B	.34422	.35693	.35831	.36571
C	.34504	.35795	.35909	.36657
D	.34702	.35980	.36119	.36873
E	.34947	.36244	.36383	.37147
b	.34992	.36293	.36432	.37199
F	.35154	.36464	.36606	.37379
G	.35366	.36694	.36842	.37621
G	.35530	.36876	.37027	.37814
H	.35606	.36955	.37102	.37892
H	.35718	.37080	.37231	.38023
H	.35850	.37224	.37378	.38180
	21.69%	22.78%	26.58%	
A	1.36561	1.36789	1.37475	
a	.36660	.36878	.37568	
B	.36737	.36956	.37650	
C	.36821	.37041	.37736	
D	.37046	.37540	.38251	
E	.37319	.37540	.38251	
b	.37374	.37590	.38305	
F	.37558	.37777	.38500	
G	.37801	.38025	.38752	
G	.37991	.38224	.38965	
H	.38075	.38310	.39052	
H	.38208	.38445	.39192	
H	.38364	.38615	.39365	

s.l. = spectral lines.

Bedson and Williams, 1881

%	n		
	H_α	H_β	H_γ
		20°	
12.02	1.35208	1.35881	1.36244
13.625	.35488	.36711	-
16.88	.35728	.36422	1.36805
100	.54095	.55384	.56128

Le Blanc, 1889

%	n_D
20°	
0	1.33325
4.79	.34153
5.31	.34260
24.13	.37635

Wegner, 1889

%	n		
	Li	H_α	D
		20°	
7.9879	1.344611	1.344996	1.346918
16.5657	.359634	.360156	.362354
21.3677	.368229	.368813	.370927
22.4036	.370169	.370766	.373305
23.6326	.372561	.372940	.375162
24.7345	.374594	.375013	.377329
	H_β	H_γ	
7.9879	1.351307	1.354613	
16.5657	.366976	.370933	
21.3677	.376045	.379824	
22.4036	.378041	.381939	
23.6326	.380270	-	
24.7345	.382463	1.386533	

Barbier et Roux, 1890

g/l	t	B
30	23.5	0.356
40	23.4	.358
50	23.5	.363
80	22.7	.372
100	23.1	.379
150	23.6	.393
200	24.2	.408
311.5	23.8	.440

$$B = \text{dispersive power} = \frac{n_1 - n_2}{1/\lambda_1^2 - 1/\lambda_2^2}$$

Schütt, 1890

%	n K	n Li	n H_α
		18°	
25	1.373228	1.375573	1.375995
20	.364120	.366417	.366813
15	.355284	.357478	.357878
10	.346535	.348676	.349042
5	.337942	.339998	.340362
3	.334517	.336537	.336895
2	.332807	.334821	.335160
1	.331073	.333081	.333430
0.5	.330727	.332207	.332557
0.3	.329871	.331862	.332199
0.1	.329520	.331502	.331860
0	.329344	.331338	.331685
%	**D**	**Tl**	**H_β**
25	1.378270	1.380707	1.383599
20	.368995	.371306	.374054
15	.359957	.362178	.364792
10	.351053	.353164	.355639
5	.342280	.344286	.346649
3	.338774	.340757	.343054
2	.337045	.338998	.341267
1	.335279	.337217	.339476
0.5	.334420	.336316	.338569
0.3	.334055	.335963	.338209
0.1	.333691	.335616	.337854
0	.333522	.335432	.337675

%	H_γ	%	H_γ
25	1.387844	2	1.344572
20	.378092	1	.342738
15	.368612	0.5	.341842
10	.359267	0.3	.341458
5	.350063	0.1	.341079
3	.346405	0	.340908

Bender, 1890

M	n H_α	t	n D	t
		15°		
0.5	1.33594	20.7	1.33776	21.0
1.0	.34085	19.6	.34270	20.3
1.5	.34561	18.1	.34756	18.9
2.0	.35024	17.6	.35219	17.6
2.5	.35472	17.9	.35663	18.6
3.0	.35874	18.0	.36100	18.1
3.5	.36294	18.9	.36504	18.1
4.0	.36691	18.4	.36916	17.8
4.5	.37116	18.3	.37329	17.3
	H_β	**t**	**H_γ**	**t**
0.5	1.34203	21.0	1.34530	20.9
1.0	.34715	19.0	.35051	20.0
1.5	.35209	18.5	.35553	18.8
2.0	.35690	17.9	.36031	18.0
2.5	.36150	18.2	.36487	18.1
3.0	.36588	18.1	.36941	18.6
3.5	.37022	18.4	.37407	17.8
4.0	.37429	17.8	.37819	18.2
4.5	.37852	18.0	.38265	17.6

Buchkremer, 1890

%	n_D	%	n_D
		20°	
0	1.33313	4.0140	1.34017
1.396	.33555	13.148	.35620
2.2625	.33715	23.800	.37605
3.018	.33835		

Borgesius, 1895

%	t	Dn_{Na} (sol.-aq.)
0.43	17.4	0.000083
0.89	17.7	.000165
1.79	18.1	.000327
3.59	19.8	.000646
6.77	19.4	.001277
12.66	18.8	.002554
12.67	18.0	.002549
22.19	18.0	.005026
52.03	19.2	.018903

Boguski, 1899

%	n_D	%	n_D
		20°	
44.89	1.38533	16.22	1.35222
39.71	.37963	9.998	.34509
34.35	.37351	4.94	.33914
29.41	.36762	2.63	.33649
25.42	.36293	0.00	.33336
20.11	.35679		

Bender, 1902

t	n H_α	n D	n H_β	n H_γ
		2.87%		
10	1.336893	1.338860	1.34029	1.346451
15	.336521	.338427	.342665	.346084
20	.336078	.338001	.342215	.345568
25	.335608	.337447	.341651	.345022
30	.335023	.336906	.341070	.344441
35	.334377	.336906	.340454	.343836
40	.333636	.335486	.339650	.343054
		5.62%		
10	1.341747	1.343717	1.348063	1.351559
15	.341287	.343264	.347615	.351050
20	.340819	.342752	.347117	.350553
25	.340277	.342200	.346523	.350032
30	.339625	.341591	.345247	.349362
35	.339029	.340946	.345247	.348756
40	.338308	.340235	.344569	.348030
45	.337593	.339485	.343750	.347220
50	.336723	.338665	.342887	.346348
55	.335900	.337826	.342021	.345471
60	.335062	.336983	.341235	.344616
65	.334206	.336150	.340389	.343826
70	.333324	.335268	.339507	.342944

10.85%

10	1.351008	1.353030	1.357604	1.361347
15	.350439	.352427	.356995	.360723
20	.349821	.351840	.356414	.360062
25	.349145	.351149	.355752	.359374
30	.348499	.350455	.355005	.358702
35	.347773	.349772	.354303	.357986
40	.347072	.349092	.353661	.357360
45	.346314	.348319	.352941	-
50	.345464	.347420	.352019	1.355606
55	.344565	.346587	.351100	.354681
60	.343719	.345703	.350224	-
65	.342810	.344729	.349273	1.352840
70	.341759	.343776	.348217	.351827

15.72%

10	1.359750	1.361904	1.366698	1.370612
15	.359077	.361201	.366025	.369904
20	.358391	.360517	.365346	.369219
25	.357665	.359787	.364618	.368486
30	.356949	.359071	.363892	.367866
35	.356249	.358357	.363170	.367169
40	.355498	.357546	.362345	.366256
45	.354691	.356780	.361532	.365422
50	.353871	.356028	.360765	.364648
55	.353016	.355122	.359830	.363729
60	.352205	.354256	.358928	.362832
65	.351319	.353409	.358080	.361965
70	.350262	.352329	.357024	.360924

20.31%

10	1.367955	1.370206	1.375210	1.379287
15	.367309	.369465	.374542	.378591
20	.366587	.368734	.373784	.377877
25	-	.367992	.373012	.377153
30	1.365100	.367279	.372303	.376413
35	.364342	.366529	.371518	.375603
40	.363576	.365707	.370749	.374790
45	.362756	-	.369949	-
50	.361924	1.364083	.369072	1.373156
55	.361093	.363226	.368236	.372333
60	.360222	.362397	.367360	.371474
65	.359361	.361510	.366494	.370572
70	.358527	.360606	.365617	.369668

22.50%

10	1.371685	1.373866	1.379099	1.383243
15	.370961	.373205	.378378	.382549
20	.370239	.372474	.377577	.381757
25	.369453	.370905	.376069	.381012
30	.368630	.370905	.376069	.380216
35	.367923	.370209	.375248	.379452
40	.367130	.369401	.374463	.378632
45	-	.368607	.373737	.377800
50	1.365523	.367764	.372826	.376940
55	.364711	.366935	.372017	.376133
60	.363876	.366102	.371162	.375256
65	.363035	.365249	.370281	.374413
70	.362200	.364380	.369379	.373545

Wagner, 1903

c	n_D	c	n_D
17.5°			
0	1.33320	9.893	1.34947
0.217	.33358	10.131	.34984
.435	.33397	.369	.35021
.657	.33435	.609	.35058
.880	.33474	.849	.35095
.104	.33513	11.089	.35132
.328	.33551	.329	.35169
.552	.33590	.570	.35205
.776	.33628	.810	.35242
2.001	.33667	12.050	.35279
.226	.33705	.290	.35316
.451	.33743	.530	.35352
.676	.33781	.770	.35388
.901	.33820	13.013	.35425
3.126	.33858	.255	.35461
.351	.33896	.498	.35497
.581	.33934	.740	.35533
.811	.33972	.983	.35569
4.041	.34010	14.225	.35606
.271	.34048	.468	.35642
.501	.34086	.710	.35678
.731	.34124	.953	.35714
.961	.34162	15.195	.35750
5.191	.34199	.440	.35786
.421	.34237	.684	.35822
.651	.34275	.929	.35858
.885	.34313	16.174	.35894
6.119	.34350	.418	.35930
.352	.34388	.663	.35966
.586	.34426	.908	.36002
.820	.34463	17.152	.36038
7.054	.34500	.397	.36074
.288	.34537	.642	.36109
.522	.34575	.890	.36145
.755	.34612	18.138	.36181
.989	.34650	.387	.36217
8.227	.34687	.635	.36252
.463	.34724	.884	.36287
.703	.34761	19.132	.36329
.941	.34798	.380	.36359
9.179	.34836	.629	.36394
.417	.34873	.877	.36429
.655	.34910	20.125	.36464

Jones and Getman, 1904

M	n_D
0°	
0.2	1.32703
0.5	.32997
1.0	.33469
2.0	.34389
3.0	.35249

Dinkhauser, 1905

%	n_D	%	n_D
18°			
0	1.33345	10.856	1.35224
2.870	.33851	15.729	.36105
4.262	.34088	20.313	.36926
5.629	.34320	24.644	.37710

Briner, 1906

c	n_D	c	n_D
	18°		
4.264	1.3454	6.667	1.3642
5.000	.3495	6.894	.3672
5.555	.3533	6.923	.3677
5.967	.3566	7.143	.3711
6.000	.3570	7.333	.3745
6.364	.3607	7.466	.3771

Chéneveau, 1907

%	n_D	%	n_D
	15°		
25.65	1.3799	14.07	1.3583
23.46	.3758	11.36	.3534
21.21	.3714	8.70	.3488
18.90	.3670	5.92	.3439
16.46	.3627	2.99	.3386

Heydweiller, 1909

M	n_D
18°	
0.5	1.33834
1.0	.34318
2.0	.35241
4.0	.36959

$Dn(H_\gamma - H_\alpha).10^3 = 36$

Guerdjikova, 1910

%	n_D	%	n_D
	25°		
0	1.33255	18.300	1.3648
11.410	.3524	24.292	.3758

Baxter, Boylston and al., 1911

%	n_D	%	n_D
	25°		
0.5280	1.33342	1.1068	1.33438
.5493	.33341	5.3562	.34169
.9980	.33417	5.4131	.34179
1.0618	.33433	14.344	.35747

Grufki, 1913

M	%	n H_α	n H_β	n H_γ
		18°		
0	0	1.33139	1.33734	1.34053
0.50	2.89	.33648	.34261	.34591
1.012	5.67	.34132	.34764	.35102
2.009	10.89	.35045	.35708	.36068
4.003	20.32	.36736	.37455	.37862

Lübben, 1913

%	n 2312.9 Å	n 2573.2 Å	n 2748.7 Å
		18°	
0	1.38895	1.37361	1.36654
2.89	.39646	.38026	.37277
5.67	.40373	.38660	.37885
10.89	.41749	.39871	.39017
20.32	.44305	.42105	.41139

%	3133.0 Å	3403.6 Å	3611.9 Å	5086.0 Å
0	1.35622	1.35062	1.34754	1.33633
2.89	.36204	.35618	.35299	.34144
5.67	.36758	.36154	.35829	.34634
10.89	.37813	.37159	.36811	.35553
20.32	.39743	.39032	.38643	.37269

Shippy and Burrows, 1918

%	n_D
25°	
5	1.34172
10	.35021
15	.35893
20	.36829

Pulvermacher, 1920

M	n_D	M	n_D
	25°		
5.149	1.3781	2.574	1.3569
4.505	.3728	1.278	.3454
3.862	.3675	0.257	.3357

Wasastjerna, 1920

M	n Hc	n D	n Hf
		18°	
0.0000	1.33168	1.33348	1.33764
0.1000	.3270	.33450	.33870
0.2000	.3367	.33552	.33975
0.5000	.3667	.33860	.34286
0.7000	.3858	.34052	.34485
1.0000	.4150	.34337	.34779
1.5000	.4608	.34801	.35257"
2.0000	.5058	.35259	.36174
2.5000	.5491	.35699	.36617
3.0000	.5911	.36123	.37054
3.5000	.6339	.36560	.37475
4.0000	.6740	.36961	.37894
4.5000	.7154	.37374	.37894
5.0000	.7546	.37768	.38308
		25°	
0.0000	1.33108	1.33291	1.33701
0.1000	.33220	.33390	.33814
0.2000	.33306	.33493	.33906
0.5000	.33595	.33788	.34211
0.7000	.33784	.33980	.34404
1.0000	.34065	.34259	.34695
1.5000	.34528	.34718	.35172
2.0000	.34963	.35165	.35637
2.5000	.35393	.35599	.36077
3.0000	.35809	.36025	.36508
3.5000	.36230	.36452	.36949
4.5000	.37042	.37272	.37784
5.0000	.37428	.37651	.38184

Flöttmann, 1928

%	t	n_D
26.343	15	1.38099
26.403	20	1.38038
26.47	25	1.37964

Geffcken, 1929

M	n_D	M	n_D
	25°		
0	1.33253	3.0568	1.35893
1.9088	.34993	3.9899	.36570
1.9886	.35058	4.9515	.37212
2.0175	.35083	4.9962	.37242
2.9997	.35850	6.0551	.37502

Herz, 1930

%	n_D	%	n_D
	18°		
2.870	1.33851	15.729	1.36105
4.262	1.34088	20.313	1.36926
10.856	1.35224	24.644	1.37710

Steyer, 1931

%	n_C	n_D	%	n_C	n_D
		20°			
1.0	1.33281	1.33468	10.0	1.34843	1.35055
2.5	.33540	.33733	15.0	.35711	.35922
5.0	.33976	.34170	20.0	.36613	.36835
7.5	.34418	.34607	25.0	.37543	.37771

Hölemann and Kohner, 1931

m	n_D	m	n_D
	25°		
0	1.33266	3.4476	1.36194
1.6692	.34804	3.4477	.36196
1.6693	.34804	5.4140	.37525
1.6693	.34805		
	35°		45°
0	1.33142	0	1.32995
1.6689	.34655	1.6689	.34497
1.6690	.34655	1.6690	.34498
3.4494	.36040	3.4497	.35872
5.4144	.37356	3.4498	.35870
		5.4186	.37185

Spacu and Popper, 1934

%	n_{He}	%	n_{He}
	20°		
25.838	1.378818	11.682	1.352953
24.850	.377000	0	.3324865
17.571	.363484		

Akhumov and Golovkov, 1935

%	n_D	%	n_D
	25°		
0	1.3334	15.70	1.3637
2.70	.3409	20.30	.3721
5.60	.3457	26.35	.3826
10.70	.3550	100	.5442

Usanovich and Mun, 1955

%	n_D	%	n_D
	25°		
5.77	1.3428	19.02	1.3661
10.39	.3510	25.00	.3770
14.57	.3582		

Electromagnetic properties

Kohlrausch and Grotrian, 1875

%	κ 0°	κ 18°	α.10⁴	β.10⁶
5	427	668	292	110
10	775	1203	290	102
15	1061	1632	279	110
20	1250	1945	290	108
24	1239	2104	311	111

$\kappa_t = \kappa_0 (1 + \alpha t + \beta t^2)$

Grotrian, 1876

%	κ	τ.10⁴
	18°	
4.97	664	213
15.00	1632	207
23.86	2098	219

Berggren, 1877

%	t	κ
7.36	9.7	700
9.03	10.0	879
15.00	10.5	1351

Kohlrausch, 1879

%	τ.10⁴	κ	%	τ.10⁴	κ
		18°			
5	218	666	25	228	2122
10	215	1203	26	231	2137
15	213	1632	26.4	234	2142
20	217				

Holland, 1898

t	τ.10⁴	κ	t	τ.10⁴	κ
		26.4 %			
10	221	1638.6	50	248	3568.3
18	229	1990.1	60	251	4087.0
20	229	2083.1	70	254	4616.2
30	229	2548.9	80	256	5146.4
40	242	3051.1			

Jones and Getman, 1904

M	λ	M	λ
	0°		
0.2	51.84	2.0	43.14
0.5	48.62	3.0	36.59
1.0	47.16		

Hechler, 1904

t	κ	t	κ	t	κ
		22.3 %			
-17.70	706	-11.29	895	+ 2.54	1378
-17.36	722	- 8.57	984	+ 4.20	1440
-16.05	765	- 6.39	1056	+ 7.59	1572
-13.96	818	- 4.04	1136	+ 9.10	1635
-12.81	853	+ 0.05	1281		

Demolis, 1906

t	κ 76.7 g/1	κ 150.5 g/1	κ 228.7 g/1	κ 295.5 g/1
20	0.0972	0.1617	0.2048	0.2245
30	0.1181	0.1975	0.2489	0.2776
40	0.1397	0.2332	0.2952	0.3306
50	0.1614	0.2688	0.3427	0.3847
60	0.1829	0.3050	0.3906	0.4403
70	0.2055	0.3424	0.4374	0.4958
80	0.2262	0.3796	-	-

g/1	κ 18°	κ 25°	κ 50°	κ 75°
75	0.0918	0.1058	0.1588	0.2123
100	0.1143	0.1327	0.1985	0.2643
125	0.1353	0.1572	0.2356	0.3153
150	0.1543	0.1789	0:2685	0.3401
175	0.1702	0.1971	0.2958	0.3729
200	0.1839	0.2126	0.3196	0.4056
225	0.1948	0.2253	0.3400	0.4384
250	0.2033	0.2356	0.3576	0.4808
275	0.2099	0.2445	0.3714	0.5045
300	0.2148	0.2525	0.3870	0.5279

Inclan, 1906

t	λ	t	λ
	5.34 mol %		
2	39.994	60.5	108.475
10	50.126	71.5	124.966
20.5	62.726	81.1	152.749
33.4	76.092	89	165.025
43.8	85.309	96	181.007
52	96.325		
	3.63 mol %		
1.5	44.031	51.2	104.499
10	56.644	61.6	125.65
19	67.598	70.5	140.42
34	90.631	79	153.43
42	95.887	88	175.29
	2.75 mol %		
1	46.648	58.1	127.268
9.8	57.959	70.5	153.33
19.2	69.070	79	159.80
31.2	85.209	87:2	181.14
40.5	100.867	97	200.57
49.5	110.370		

Clausen, 1912

M	λ	τ.
	0°	
0.884	48.64	430
1.830	42.56	779
2.843	37.50	1066
3.924	32.00	1255
5.085	26.16	1330
5.325	25.16	1340

Mc Kie, 1918

M	κ	M	κ
	18°		
5.013	2128	0.5979	477.3
3.010	1694	0.4781	389.5
2.005	1288	0.3360	280.1
1.994	1287	0.2004	173.1
0.9986	736.7	0.01002	10.24
0.7964	612.6		

Giordani and Maresca, 1929

t	κ	t	κ
	20°		
N/2		N/1	
19	411	20.1	776
35.6	563	35.4	1029
44.9	646	44.8	1198
65.2	849	65.4	1570
150 g/1		250 g/1	
12.08	1354	12.5	1782
35.95	2143	34.9	2790
45.35	2487	45.3	3304
65.40	3145	63.2	4188
300 g/1		300 g/1	
11.70	1839	45.25	3528
35.10	2985	61.20	4400

Desai, Naik and Desai, 1934

M	λ
30°	
4.0	45.20
3.0	54.70
2.0	58.70
1.0	70.03
0.5	81.80
0.25	99.70
0.02	130.95
0.005	134.60

Rodnyanskii and Galimker, 1955 (fig.)

t	λ		
	1 N	2 N	3 N
25	86.4	75.9	66.0
50	143	122	106
100	236	202	178
150	318	272	239
200	392	334	292
250	440	368	322
300	448	365	315
340	419	341	290

Gellings, 1956

N	λ	N	λ
	25°		
0.0089148	114.42	1.8886	64.14
.044921	106.59	2.4813	57.91
.18611	95.89	3.0487	52.30
.46425	86.33	3.4976	48.97
.69031	81.27	4.0068	45.17
.95575	76.77	4.4912	41.83
1.1612	73.59	4.8497	39.55
.6157	67.32		

Chambers, Stokes and Stokes, 1956

N	λ	N	λ
	25°		
method I		method II	
0.15621	103.66	0.11300	105.91
.20998	101.36	.12452	105.24
.24584	100.02	.18776	102.19
.33998	97.29	.26492	99.41
.49935	93.66	.36091	96.71
.68703	90.26	.46914	94.25
1.05142	85.12	.60766	91.60
.3948	81.02	.66232	90.65
.5193	79.65	.74528	89.34
.9279	75.41	.86366	87.58
2.3793	71.14	1.2312	82.87
.7439	67.82	.4647	80.215
3.1431	64.36	.8416	76.277
.5037	61.30	2.6739	68.445
4.5199	53.13	3.1723	64.088
5.3540	46.83	.8769	58.228
		4.3281	54.622
		.8114	50.882

Fink, 1885

P	κ 0°	κ 18°
	0.99%	
1	119.9	193.7
109	121.8	195.3
200	123.1	196.9
300	125.0	198.5
400	126.5	199.9
500	128.0	201.1
	4.99%	
1	484.2	778
109	490.0	783
200	495.0	787
300	500.0	790
400	504.5	797
500	509.0	801
	9.90%	
1	872	1388
109	880	1395
200	887	1400
300	891	1406
400	898	1412
500	902	1417
	14.82%	
1	1216	1865
109	1221	1873
200	1227	1878
300	1230	1884
400	1234	1888
500	1237	1891
	19.60%	
1	1417	2174
500	1428	2181
	24.90%	
1	1518	2488
500	1509	2474
	26.39%	
1	1572	2528
500	1557	2558

Williamson, 1922

P	R/R_1	P	R/R_1	P	R/R_1
0.292%	25°	20.259%	30°	23.983%	25°
1	1.0000	1	1.0000	1	1.0000
492	0.96636	247	0.99922	262	1.0023
987	.94552	490	1.0011	500	.0055
1507	.93410	751	.0038	760	.0100
1861	.93100	1025	.0112	1010	.0153
2106	.93048	1506	.0227	1245	.0211
2380	.93158	2036	.0334	1513	.0285
2653	.93416	2445	.0534	2085	.0473
3435	.94826	3071	.0689	2531	.0645
4120	.96796	3490	.0869	3123	.0903
3070	.94028	2816	.0449	3500	.1082
1949	.93006	1827	.0174	3872	.1267
0.99	.99984	872	.0017	2636	.0686
				1707	.0340
				815	.0105
				374	.0032

Adams and Hall, 1931

P kg	R	P kg	R
	20.26 % at 29.93°		
1	64.31	3207	67.98
245	64.26	3523	68.74
502	64.28	3979	69.90
769	64.38	2842	67.20
1019	64.53	1844	65.43
1546	65.06	880	64.42
2073	65.80	1	64.31
2475	66.52		
	23.79 % at 24.93°		
1	90.963	1066	92.344
300	91.181	1308	92.853
538	91.463	1571	93.475
807	91.868		
	23.92 % at 29.93°		
1	59.85	2021	62.26
246	59.98	2574	63.26
505	60.17	3149	64.65
779	60.43	3585	65.74
1008	60.69	3969	66.78
1500	61.37		
	23.98 % at 24.93°		
1	65.92	3228	71.87
271	66.07	3613	73.05
517	66.28	4001	74.27
786	66.58	2726	70.44
1046	66.93	1769	68.16
1290	67.31	832	66.61
1568	67.80	386	66.13
2158	69.04	1	65.92
2617	70.17		
	26.09 % at 0°		
1	132.26	2933	150.88
291	133.05	3353	154.76
487	133.83	2443	146.73
1007	136.26	1387	138.84
1533	139.34	487	134.19
1940	142.19	1	132.97
2338	145.35		
	26.48 % at 29.93°		
1	58.19	2554	62.52
234	58.42	3169	64.13
509	58.74	3491	64.93
769	59.08	3942	66.21
1011	59.45	2613	62.64
1259	59.86	1610	60.49
1532	60.34	745	59.03
2169	61.65	1	58.20

Dewar and Fleming, 1897

t	ε	t	ε
	23.6%		
- 203.1	2.78	-165.2	4.73
- 200.0	2.62	-156.5	5.88
- 192.0	2.90	-122.7	27.40
- 177.0	3.68	-108.3	156.00

Perkin, 1894

%	(α) magn.
15.5°	
26.174	1.3769

Oppenheimer, 1898

%	(α) magn.
20°	
20.22	1.64
15.56	1.66
10.90	1.67
5.49	1.66

Agerer, 1905

%	t	(α) magn.
4.889	18.1	1.597
22.639	18.4	1.644

Guerdjikova, 1910

%	(α) mol magn.	%	(α) mol magn.
	25°		
0	5.068	18.300	6.132
11.410	5.647	24.292	6.564

Mc Clung and Mc Intosh, 1902

d	% X-ray absorption
at room t.	
1.153	84.2
.078	76.1
.039	67.0
.000	63.0

Okazaki, 1933

%	Verdet's constant.10^5 (3441 Å)
28°	
4.44	4698
9.47	5057
16.10	5542
19.89	5813
22.85	6076
23.39	6126
24.89	6294

Okazaki, 1936

%	Verdet's constant.10^5 (D)
25°	
0	1307.5
5.314	1402
8.131	1458
12.70	1537
16.66	1616
20.77	1697
24.73	1781
26.06	1806

Okazaki, 1951

%	Verdet's constant.10^5 (D)
35°	
0	1304.7
10.35	1478
13.87	1538
17.03	1601
21.39	1685
25.02	1762

Heat constants.

Person, 1851

%	U
24.58	0.7852
21.53	.8014
12.07	.8721
6.34	.9288

Thomsen, 1871

mol%	U
18°	
9.1	0.791
4.8	.863
3.2	.895
2.0	.931
1.0	.962
0.5	.978

Schüller, 1869

%	U	%	U
4.76	0.9306	10.00	0.8079
9.09	.8909	23.08	.7897
13.04	.8606	24.15	.7752
16.67	.8304	25.93	.7713

Lundquist, 1869

%	t	U
0	40.3	1
23.43	43.9	0.785

Winkelmann, 1873

%	U	%	U
3.09	0.9638	17.12	0.8526
5.15	.9449	26.03	.8072
11.05	.8925		

Blümcke, 1884

%	t	U
4.76	16.5	0.958
9.12	15	0.912
13.13	15.3	0.877
16.82	14.7	0.852
23.02	17	0.811

Tendt, 1899

t	U	t	U
18.81%		10.38%	
34.7	0.8412	31.8	0.8925
35.2	.8431	32.8	.8948
35.0	.8401	37.0	.9017
41.7	.8511	43.4	.9088
41.9	.8500	46.0	.9076
43.0	.8513	50.1	.9144
44.6	.8492	52.5	.9120
45.9	.8545		
50.3	.8569	0%	
50.8	.8563	34.1	0.9962
50.8	.8563	38.1	1.0076
51.5	.8536	42.2	.0082
51.5	.8547	46.1	.0170
54.5	.8539	50.1	.0230

Demolis, 1906

g/l	U	g/l	U
50	0.9366	200	0.8321
75	0.9171	225	0.8185
100	0.8982	250	0.8060
125	0.8802	275	0.7950
150	0.8631	300	0.7840
175	0.8471		

Gröber, 1908

%	U -20°	-10°	0°	+10°	+20°
0	-	-	1.006	1.002	0.999
2	-	-	0.966	0.969	0.973
4	-	-	0.944	0.947	0.951
6	-	-	0.923	0.927	0.930
8	-	-	0.904	0.907	0.910
10	-	-	0.885	0.889	0.892
12	-	-	0.869	0.872	0.875
14	-	0.851	0.854	0.857	0.860
16	-	0.838	0.840	0.843	0.845
18	-	0.825	0.827	0.830	0.832
20	-	0.812	0.814	0.816	0.818
22	0.795	0.798	0.801	0.804	0.806
24	-	0.787	0.789	0.791	0.793
26	-	0.776	0.778	0.780	0.781

%	+30°	+40°	+50°	+60°
0	0.998	0.999	1.001	1.003
2	0.977	0.980	0.983	0.986
4	0.955	0.958	0.961	0.964
6	0.933	0.936	0.939	0.942
8	0.913	0.916	0.919	0.922
10	0.895	0.898	0.901	0.903
12	0.878	0.880	0.883	0.885
14	0.862	0.864	0.867	0.870
16	0.848	0.850	0.853	0.855
18	0.834	0.836	0.839	0.841
20	0.821	0.823	0.824	0.827
22	0.808	0.810	0.812	0.814
24	0.795	0.797	0.799	0.801
26	0.783	0.785	0.787	0.788

Bousfield, 1919

%	U 7°	20°	33°
25.000	0.787	0.7875	0.788
18.945	.820	.8225	.825
13.519	.858	.862	.8655
10.005	.886	.890	.893
5.625	.931	.935	.938
2.8669	.962	.963	.9645
1.4476	.981	.982	.9825
0	1.004	1.008	1.008

Gerlach, 1926

%	U	t. limits	%	U	t. limits
0.8	0.9889	20	12.3	0.871	18
1.6	0.978	18	14.0	0.863	18
3.1	0.9605	20	-	0.870	60
3.2	0.962	-	-	0.851	-10
4.9	0.945	19-46	18.8	0.841	17-52
6.1	0.931	19	-	0.854	20-89
9.8	0.895	19	24.3	0.792	18-20
10.3	0.892	15-49	24.5	0.791	18
-	0.912	15-90	26.0	0.776	-10
11.5	0.8779	20	-	0.788	60
-	0.877	16.52			

Randall and Rossini, 1929

m	U	m	U
	25°		
0.00	0.9979	0.50	0.9629
.01	.9971	.75	.9478
.02	.9963	2.00	.9338
.05	.9940	2.25	.9209
.10	.9902	1.50	.9088
.20	.9829	2.00	.8872
.35	.9726	2.50	.8684

Nikolaev, Kogan and Ogorodnikov,1936

%	U	%	U
	25°		
1.94	0.9760	19.41	0.8221
7.02	.9224	20.04	.8200
14.10	.8618	25.02	.7865

Young and Machin, 1936

m	U (by mol)		
	0.0°	12.5°	25.0°
5.9536	-410.5	-234.6	- 80.41
	-411.0	-233.9	- 80.57
5.4270	-489.0	-304.1	-144.4
	-490.2	-304.7	-145.5
4.4580	-618.8	-417.8	-249.5
	-618.8	-417.7	-249.7
3.5133	-733.1	-510.4	-333.3
	-733.8	-510.1	-331.8
2.4504	-832.6	-576.4	-388.9
	-833.1	-576.0	-389.3
1.5023	-862.5	-581.4	-380.0
	-861.1	-579.4	-380.3
0.77721	-768.6	-478.9	-386.1
	-767.4	-479.5	-
0.49730	-643.3	-374.5	-199.1
	-644.4	-378.3	-198.5
0.19518	-369	-153	- 23
	-362	-150	- 14
0.09305	-164	-	-
	-207	-	-
	-142	-	-

Rutskov, 1948

%	U		
	25°	50°	75°
1.156	-	0.9840	0.9880
2.840	-	.9645	.9686
6.810	-	.9236	.9355
12.75	-	.8710	.8695
24.50	0.790	.787	.780

Winkelmann, 1873

%	Q diss. at 0°	Q diss. at 50°
3.09	-27.12	-6.00
5.15	24.20	6.33
7.98	22.05	-
11.05	20.19	5.49
14.06	18.29	-
17.12	16.30	5.22
20.01	14.88	-
23.01	13.52	-
26.03	11.95	4.61
29.03	10.69	-
32.04	9.59	-

Scholz, 1892

M	Q diss. (by gr.NaCl)
at room temperature	
0.0625	-33.07
.125	33.25
.25	32.62
.5	30.60
1.0	27.13
2.0	21.02
4.0	13.18
6.0	9.00

Staub, 1890

%	Q diss.	%	Q diss. (by gr.NaCl)
at room temperature			
2.45	-28.09	20.63	-11.91
6.36	23.79	23.08	10.03
8.50	21.51	23.10	10.17
11.50	18.61	25.50	8.95
14.60	16.08	26.15	8.10
16.00	15.14	26.27	8.22
20,31	12.25		

Colson, 1901

t	Q diss. (by gr. NaCl)	
	25%	75%
17.5	-21	-17.7
28.6	-15.8	-13.9
36.5	-12.5	-11.4
101	+10	+6.7

t	D t	t	D t
31 g/100 cc + 400 cc H_2O			
15	-0.715	46.6	-0.085
36.7	-0.25	53.5	+0.03
43.8	-0.15	54.2	+0.03

Varali-Thevenet, 1902

%	Q diss. (by gr. NaCl) 1	2
0.2	-33.8	-
.4	31.8	-
.6	29.2	-
.8	29.2	-
1.0	28.4	-
2.0	27.1	-29.6
4.0	23.7	23.1
6.0	19.9	19.0
8.0	16.6	15.8
10.0	13.3	12.4
12.0	9.3	10.5
14.0	6.4	7.4
16.0	5.2	4.7
18.0	3.2	3.8

Stearn and Smith, 1920

m	% initial	+moles aq.	Q dil. (by mole aq.)
3.2	10.586	27.4	-19.9
		18.45	21.6
		11.22	22.95
1.6	8.280	34.83	7.04
		22.06	8.06
		11.75	8.67
		7.83	9.08
0.8	7.986	36.45	1.92
		32.86	2.37
		25.65	2.76
		16.75	2.89
0.4	7.900	37.72	0.79
		25.58	0.94
		21.5	1.18
		16.5	1.28
		12.47	1.38
		8.86	1.41
0.2	9.300	28.8	0.61
		23.4	0.765
		18.11	0.74
		14.6	-0.69

Lehtonen, 1922

M	Q diss. (by gr. NaCl)
0°	
0.0625	-33.517
.1250	-33.174
.2500	-32.256
.5000	-30.309
1.0000	-26.908
2.0000	-20.898
4.0001	-13.106

Lipsett, Johnson and Maass, 1927

%	Q diss.
20°	
0.22	-1162
1.200	-1158
1.200	-1152
4.620	-1042.8
4.620	-1043.6
8.085	-915.5
8.085	-914.6
10.732	-819.3
10.732	-818.9
13.946	-710.7
13.946	-710.9
19.540	-560.2
19.540	-560.0
25.791	-460.8
25.791	-460.6
25°	
0.3984	-1004
0.3953	-1015
1.3831	-1006
1.9631	-1000
3.0892	-970.6
5.6700	-896.3
8.0855	-819.4
10.732	-737.3
13.799	-651.5
16.431	-586.3
19.372	-524.6
24.321	-457.5

Backlund, 1911-12

mol%	Q dil.
1.32	-82
.58	-118
.95	-163
2.60	-236
3.25	-309
3.84	-374
2.68	-482
6.25	-590
7.32	-665

Q dil. = heat of dilution of resp. solution to 1 mole NaCl in 150 moles H_2O

Next Page

Vüst and Lange, 1925

mol %	Q diss. integral	mol %	Q diss. intermediate
0.925	-982	0.4668	-982
1.655	-919	1.285	-840
2.367	-854	2.010	-705
3.846	-725	3.104	-526
5.244	-623	4.564	-359
7.251	-517	6.263	-258
9.837	-459	8.236	-262
0.758	-999	0.3830	-999
1.533	-932	1.146	-864
2:282	-863	1:920	-725
3.741	-734	3.005	-539
5.142	-631	4.456	-373
7.191	-519	6.232	-260
10.086	-458	8.210	-257

Chipman, Johnson and Maass, 1928, 1929

M	Q diss.	M	Q diss.
0.202	-1193	1.63	-836
.404	-1159	2.12	-755
.644	-1023	2.62	-686
.821	-1010	3.18	-560
1.23	-918	4.03	-477

Lundquist, 1896

%	t	heat conductivity
0	40.3	0.0933
23.43	43.9	.0895

Beetz, 1879

d (20°)	heat conductivity	t. limits
1.200	433	8 - 14°
"	649	36 - 28°
1.110	427	8 - 14°
"	668	36 - 28°
aq.	413	8 - 14°
"	662	36 - 28°

Jäger, 1891

%	heat conductivity coeff. 1	2
	at room temperature	
0	100	-
12.5	96.8	96.4
25	93.9	92.3
	(by two methods)	

Kapustinskii and Ruzavin, 1955

%	$c.10^6$	%	$c.10^6$
	25°		
4.85	1449	14.4	1424
5.00	1446	18.1	1413
9.53	1433	20.9	1402
9.75	1435	22.2	1404
13.7	1423		

c = heat conductivity coefficient

(second series).

%	$c.10^6$	%	$c.10^6$
4.15	1446	17.6	1407
8.42	1435	24.4	1392
12.9	1421		

Kapustinskii and Ruzavin, 1956.

m	H
25°	
1	0.075
1.64	0.072
2.56	0.068
3.60	0.068
5.76	0.066

$$H = \frac{1000 + m\,M/K - 1000\,Ko}{m}$$

m = molality

K = heat conductivity coefficient of the solution

Ko= " " " of the water

Previous Page

Water (H_2O) + Sodium bromide (NaBr)

Heterogeneous equilibria.

Tammann, 1885

t	p 0%	15.42%	24.20%	29.09%	36.53%
43.69	67.2	63.0	59.1	56.9	51.7
45.08	72.2	67.8	63.5	60.9	55.3
47.34	81.0	76.3	71.4	68.6	62.6
48.92	87.7	82.2	77.0	73.5	67.0
52.53	104.8	98.5	92.6	88.7	80.4
56.31	125.7	118.4	111.5	106.5	96.3
58.07	136.6	128.6	120.7	115.2	104.1
60.08	150.0	141.4	132.7	126.5	114.5
62.42	167.0	157.2	147.7	140.6	127.4
65.07	188.3	177.2	166.2	158.3	143.5
67.67	211.2	198.0	180.0	178.7	160.7
68.81	222.0	208.5	195.6	184.0	168.7
70.96	243.7	228.2	214.5	204.4	185.1
74.37	281.7	264.2	247.8	236.3	213.2
79.51	348.5	326.8	306.2	292.5	266.1
81.89	383.6	359.7	337.5	322.0	290.9
84.85	431.3	404.1	379.7	361.7	331.5
87.70	481.8	451.9	428.1	403.5	366.4
90.76	541.3	506.6	475.7	453.4	410.1
93.81	606.8	568.8	533.4	508.9	462.1
97.03	682.8	640.5	601.5	572.6	519.5
100.33	768.9	719.9	675.6	644.7	586.1

%	p	%	p	%	p
100°					
4.65	747.6	23.98	669.3	38.53	558.8
6.55	743.1	28.05	642.7	42.85	516.3
14.73	714.8	33.72	602.3	46.54	476.3
18.22	699.5				

Pearce, Taylor and Bartlett, 1928 and
Pearce and Nelson, 1932

m	p	m	p	m	p
25°					
0.0	23.752	2.0	22.128	6.0	17.603
0.1	23.675	2.5	21.670	6.5	16.934
0.2	23.608	3.0	21.181	7.0	16.234
0.4	23.448	3.5	20.655	7.5	15.570
0.6	23.292	4.0	20.094	8.0	14.888
0.8	23.136	4.5	19.501	8.5	14.231
1.0	22.973	5.0	18.876	9.0	13.582
1.5	22.556	5.5	18.242	9.131	13.409

Menzel and Sieg, 1932

t	p	t	p
sat. sol. + 2 aq.			
15.5	7.8	25.0	13.4
20.0	10.5	27.5	15.6

Lannung, 1934

m	p	m	p
18°			
1.818	14.6	4.349	12.7
1.988	14.4	5.406	11.9
2.603	14.0	6.76	10.9
2.928	13.8	-	9.3 sat.sol. (NaBr.2aq)
3.488	13.3		
		8.75	5.1 NaBr

Lescoeur, 1890

t	dissociation p (4+1)	p sat. sol.
10	3.9	5.45
15	-	7.4
20	7.65	9.75
30	15.5	16.3

Dingemans, 1935

t	p NaBr.2aq+ NaBr+V	NaBr.2aq+ sat.sol.+V	NaBr+ sat.sol.+V	L + V
10.00	2.7	5.8	-	9.2
12.00	3.1	6.5	-	10.5
15.00	4.0	7.8	-	12.8
17.00	5.0	9.3	-	15.5
20.00	5.8	10.5	-	17.5
22.00	7.3	12.3	-	21.1
25.00	8.4	13.7	-	23.8
27.00	9.7	15.2	-	26.7
30.00	12.0	17.8	-	31.8
32.00	13.9	19.8	-	35.7
35.00	17.0	23.0	-	42.2
37.00	19.5	25.3	-	47.1
40.00	23.8	29.2	-	55.3
42.00	27.2	32.1	-	61.5
45.00	33.0	36.9	-	71.9
47.00	37.5	40.3	-	79.6
50.00	45.2	45.8	-	92.5
50.60	46.4	46.9	46.9	95.3
52.00	-	-	50.2	102.1
55.00	-	-	58.1	118.0
57.00	-	-	64.1	129.8
60.00	-	-	74.0	149.4
62.00	-	-	81.0	163.8
65.00	-	-	92.7	187.5
67.00	-	-	101.3	205.0
70.00	-	-	115.8	233.7
72.0	-	-	126	254.6
75.0	-	-	143	289.1
77.0	-	-	156	314.1
80.0	-	-	176	355.1
82.0	-	-	191	384.9
85.0	-	-	216	433.6
87.0	-	-	233	468.7
90.0	-	-	262	525.8
92.0	-	-	282	567.0
95.0	-	-	316	633.9
97.0	-	-	340	682.1
100.0	-	-	379	760.0

Keevil, 1942

mol %	t	P
	sat.sol.	
24.9	293	35.4
32.8	396.0	92.56
39.0	460.0	131.6
43.6	493.8	153.3
55.5	567.0	174.9
78.3	678	109.7

Baroni, 1893

%	b.t.	%	b.t.
1.34	100.117	10.69	101.191
2.82	100.258	13.72	101.628
5.14	100.507	17.20	102.203
7.78	100.816		

Schlamp, 1894

%	b.t.	%	b.t.
0	100	9.96	101.048
1.33	100.12	11.65	101.272
4.12	100.388	13.61	101.565
7.14	100.710		

Johnston, 1906

%	b.t.	%	b.t.
10.39	101.050	50.24	111.240
16.56	101.722	51.11	111.902
19.38	102.182	52.24	112.845
49.12	110.300	52.65	113.100

Jablezynski and Kon, 1923

m	b.t.	m	b.t.
0.3643	100.340	1.5319	101.603
0.7444	100.730	1.7501	101.866
1.1082	101.118	1.9970	102.169
1.3324	101.372	2.2320	102.472

Robinson, 1935

m_1	m_2	m_1	m_2
0.1240	0.1226	2.248	2.002
.1368	.1351	.520	.228
.1510	.1492	.631	.335
.1611	.1591	.651	.451
.1919	.1895	.797	.460
.3570	.3470	.805	.460
.5771	.5529	3.203	.767
.6738	.6431	.492	.998
.8214	.7818	.610	3.089
.9181	.8664	.671	.137
.9587	.9014	.903	.310
1.185	1.104	4.125	.321
.385	.278	.265	.589
.491	.370	.717	.913
.891	.713	.81	.992
.975	.756		
2.144	.926		

m_1= molality of potassium chloride

m_2= " " sodium bromide

Isopiestic solutions at 25°

Robinson and Stokes, 1949

m	Osmotic coefficient	m	Osmotic coefficient
	25°		
0.1	0.934	1.2	0.969
0.2	0.928	1.4	0.983
0.3	0.928	1.6	0.997
0.4	0.929	1.8	1.012
0.5	0.933	2.0	1.028
0.6	0.937	2.5	1.067
0.7	0.942	3.0	1.107
0.8	0.947	3.5	1.150
0.9	0.953	4.0	1.199
1.0	0.958		

Kremers, 1856

%	f.t.	%	f.t.
43.67	0	52.91	80
42.92	20	53.47	100
51.02	40	sat.sol.	121 b.t.
52.63	60		

Rüdorff, 1862

%	f.t.	%	f.t.
3.68	- 1.25	10.71	- 4.25
6.03	- 2.2	20.82	-10.1

De Coppet, 1883

%	f.t.	%	f.t.
	anhydre		
53.62	44.1	54.99	86.0
53.74	51.5	54.49	90.5
53.87	55.1	54.53	97.2
53.91	60.3	54.67	100.3
53.97	64.5	55.09	110.6
54.21	74.5	55.36	114.3
54.25	80.5		
	(2+1)		
41.52	-21.3	50.26	34.6
43.37	- 6.5	51.45	40.1
44.22	0.0	51.51	40.5
44.91	+ 3.7	52.40	41.7
45.14	4.0	53.62	44.1
46.27	12.7	52.56	44.9
46.30	13.15	52.56	45.0
48.02	22.8	52.51	45.25
48.08	23.1	52.78	46.2
48.23	23.3	52.78	46.7
48.41	24.7	53.17	47.75
48.48	25.1	53.15	47.8
48.61	25.3	53.34	48.3
48.58	25.85	53.35	48.5
48.67	26.4	53.38	48.85
49.39	29.6	53.53	49.2
49.25	29.8	53.49	49.3
49.85	32.6	53.66	49.8
50.26	34.5		

Etard, 1894

%	f.t.	%	f.t.
40.1	-22	59.0	180
42.5	-10	60.9	210
56.5	+140	61.0	212
57.5	163	62.0	230
59.5	180		

Jones, 1904; Jones and Getman, 1904;
Jones and Bassett, 1905

m	f.t.	m	f.t.
0.07	-0.245	1.55	-6.200
0.13	0.462	2.07	8.610
0.26	0.907	2.59	11.350
0.52	1.842	3.10	14.000
1.03	3.815	3.62	18.000

Scott and Frazier, 1927

48.61 % f.t. = 25°

Nikolaev and Ravich, 1931

%	f.t.	
48.56	25	NaBr.2aq.
52.56	44.5	NaBr.2aq.
54.15	65	NaBr

Scatchard and Prentiss, 1933

m	f.t.	m	f.t.
0.000780	-0.0028	0.18545	-0.643
0.001558	-0.0057	0.23609	-0.818
0.002859	-0.0105	0.27647	-0.955
0.003263	-0.0121	0.33054	-1.141
0.005966	-0.0217	0.39038	-1.345
0.010280	-0.0374	0.43827	-1.509
0.010303	-0.0372	0.50805	-1.747
0.015339	-0.0553	0.56626	-1.951
0.023990	-0.0861	0.62685	-2.163
0.035356	-0.1260	0.69944	-2.415
0.054396	-0.1923	0.77676	-2.690
0.077400	-0.2718	0.87964	-3.051
0.098178	-0.344	0.96594	-3.362
0.12220	-0.426	1.0883	-3.798
0.14571	-0.507	1.2071	-4.229
0.16169	-0.561		

Distanov, 1937

f.t.	%	f.t.	%
107	54.2	199	58.1
140	55.6	222	59.5
155	56.5	225	59.7
177	57.3	228	59.8
183	57.6	232	59.9
190	57.7	242	60.7
194	57.8	248	60.8

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.	
12.05	-3.1(2+1)	16.79	47.1	0 aq.
12.89	+8.4	16.83	53.9	
14.15	24.3	16.90	72.9	
15.22	40.5	16.94	97.3	
16.64	49.0	17.89	120.1	
16.75	50.15	18.15	144.6	
16.83	50.95	19.32	171.2	

Properties of phases .

Kremers, 1858

%	d	%	d
	19.5°		
5	1.038	30	1.279
10	.078	35	.342
15	.123	40	.408
20	.172	45	.481
25	.224	50	.562

t	d			
0	1.1502	1.2914	1.3993	-
19.5	.1431	.2805	.3862	1.5071
40	.1334	.2681	.3720	.4911
60	.1222	.2550	.3576	.4752
80	.1096	.2408	.3426	.4590
100	.0955	.2263	.3271	.4425

% from last table .

Forster, 1878

%	t	d
0	24.0	0.9973
12.04	24.0	1.0982
23.61	24.8	1.2100

Wegner, 1889

%	d	%	d
	20°		
31.0222	1.296156	4.9508	1.037602
20.4406	1.178842	2.6741	1.019132
11.9564	1.097755		

Jahn, 1891

%	d
20°	
16.04	1.1350
10.29	1.0823

Humburg, 1893

%	d
16°	
0	0.9990
29.705	1.2793

Leblanc and Rohland, 1896

%	d
20°	
17.15	1.1453
22.72	1.2039

Oppenheimer, 1898

%	d	%	d
	20°		
33.57	1.3284	9.41	1.0750
25.10	1.2284	0	0.9982
17.77	1.1526		

Jones and Getman, 1904 and Jones and Bassett, 1905

%	d	%	d
	0°		
0.718	1.004576	14.25	1.120624
1.326	.009560	18.36	.161028
2.627	.019692	22.32	.199536
5.169	.038152	25.79	.238440
9.897	.082284	29,17	.278456

Loewenfeld, 1905

%	d	%	d
	15°		
0	0.9993	24.9	1.2254
9.7	1.0772	31.1	1.2951
17.6	1.1550		

Cheneveau, 1907

%	d	%	d
	22°		
21.83	1.1946	11.88	1.0990
18.05	1.1569	9.68	1.0793
14.01	1.1180	5.02	1.0399

Getmann, 1907

%	d	%	d
	18°		
9.30	1.0742	36.54	1.3681
17.40	.1489	41.67	.4401
24.52	.2229	46.46	.5066
30.87	.2958		

Heydweiller, 1909

M	d	M	d
	18°		
0.05	1.00258	1.0	1.0769
.1	.00658	2.0	.1538
.2	.01447	4.0	.3041
.5	.0381		

Baxter, Boylston and al., 1911

%	d	%	d
	25°		
0	0.99707	4.5543	1.03268
0.4857	1.00074	4.6241	.03320
0.9110	.0041	6.4288	.0490
1.1162	.0058	7.2055	.0544
1.9817	.0123	7.3936	.05600
2.1044	.01354	8.8063	.06832
2.8104	.01931	17.261	.14479
2.9218	.0198	46.06	.5005

Grufki, 1913

M	d	M	d
	18°		
0	0.99862	1.994	1.15332
0.5023	1.03824	3.975	1.30232
1.008	1.07751		

Gropp, 1915

t	d	t	d
	3.885 N		
0	1.3019	78	1.2571
18	1.2945	100	1.2414
48	1.2769		

Baxter and Wallace, 1916

%	d	%	d
70.19°		50.04°	
40.61	1.38546	40.32	1.39663
23.71	1.18338	23.50	.19418
10.59	1.06105	10.49	.07157
5.49	1.01937	5.43	.02982
2.79	0.99856	2.76	.00894
25°		0°	
39.97	1.40972	39.64	1.42228
23.29	.20608	23.12	.21601
10.47	.08200	10.33	.08854
5.39	.03962	5.37	.04444
2.73	.01833	2.72	.02222

Manchot, Jahrstorfer and Zepter, 1924

%	d	%	d
	25°		
11.878	1.0849	11.527	1.0829
22.334	1.1645	23.054	1.1668
45.405	1.3338	46.108	1.3307

Herz and Martin, 1924

t	d	t	d
	71.880 g/100 cc		
38	1.5084	50	1.4993
40	1.5070	52	1.4978
43	1.5047	54	1.4962
46	1.5023	56	1.4948
58	1.5008	58	1.4933

Hrynakowski, 1927

50.40° d = 1.6082 sat. sol.

Scott and Frazier, 1927

48.61 % (sat. sol.) d at 25° = 1.544181

Kohner, 1928

N	d	N	d
	25°		
0	0.99654	3.99924	1.26947
2.01264	1.07195	4.99933	1.32775
3.00158	1.20790	6.02882	1.38241

Scott and Durham, 1930

%	t	d
	sat. sol.	
44.47	0.00	1.4957
50.48	35.00	.5655
52.66	42.21	.5950
53.80	50.21	.6108
54.10	60.17	.6073
54.42	75.44	.5986
54.83	91.95	.5915

Shibata and Hölemann, 1931

%	d	%	d
	25.0°		
0	0.99707	42.85	1.28124
12.13	1.08330	53.48	1.34108
29.92	1.20339		
	35.0°		
0	0.99406	26.02	1.17360
15.56	1.10514	53.46	1.33434
	45.0°		
0	0.99024	25.66	1.16587
14.13	1.09068	53.48	1.32768

Scott, Obenhaus and Wilson, 1934

%	d	π
	35°	
49.295	1.54400	23.42
32.380	1.30294	29.74
20.7701	1.17443	34.03
9.8333	1.17392	38.10

Guillaume, 1946

%	d
20°	
9.46	1.0776
20.66	1.1836

Pohl, 1852 and 1860

%	π (relative)
13.9°	
0	1.000
8.57	0.899
15.60	0.818
23.59	0.764
31.27	0.665
37.00	0.651

Freyer, 1931

%	π	t	π
20°		30%	
1	45.56	15	30.80
6	42.89	20	30.64
10	40.77	25	30.56
16	37.60	30	30.53
20	35.55	40	30.56
30	30.64	50	30.67

Gibson, 1935

%	π (1-1000 atm.)
25°	
0.00	39.35
5.00	37.5
15.00	34.1
25.00	30.4
35.00	27.0
45.00	23.6

Getman, 1907

%	η
18°	
9.30	1149
17.40	1231
24.52	1347
30.87	1556
36.54	1805
41.67	2097
46.46	2591

Herz and Kuhn, 1908

and Herz and Martin, 1921

t	η	t	η
	71.880 g/ 100 cc		
38	1876	50	1371
40	1642	52	1329
43	1555	54	1288
46	1472	56	1248
48	1420	58	1209

Scatchard and Prentiss, 1933

m	κ	m	κ
	10°		
1.41830	872 36	0.10307	83.36
1.13079	707.85	0.053848	43.37
0.97107	622.32	0.035359	29.03
0.68019	455.95	0.018742	15.77
0.53910	363.33	0.007352	6.36
0.37751	267.41	0.007195	6.23
0.34046	243.95	0.003151	2.77
0.30817	222.26	0.002179	1.92
0.20002	149.23	0.0010076	0.09
0.160514	121.51		

Hrynakowski, 1927

50.40° η = 1871 σ = 80.13 sat. sol.

Loewenfeld, 1905

%	σ	%	σ
	15°		
0	74.43	24.9	75.40
9.7	78.04	31.1	78.21
17.6	79.16		

Kremers, 1857

%	t	n_D
28.88	17	1.3792
45.17	17	1.4144

Börner, 1869

t	$n_{H\alpha}$	t	$n_{H\beta}$	t	$n_{H\gamma}$
		10%			
45.9	1.340310	45.2	1.346928	44.5	1.350547
40.75	.341108	41.3	.347574	40.5	.351280
35.5	.342008	35.9	.348434	36.5	.351941
30.15	.342763	30.3	.349294	30.4	.352841
		20%			
45.3	1.351870	44.4	1.359112	43.5	1.363164
41.4	.352552	40.3	.359823	39.5	.363874
36.9	.353335	36.5	.360428	35.9	.364442
32.3	.354050	32.1	.361175	32.0	.365151
27.75	.354835	27.8	.361921	27.9	.365824
		30%			
45.1	1.362561	44.7	1.370248	44.0	1.374518
45.85	.363377	40.45	.371060	39.7	.375382
34.9	.364371	34.75	.372120	34.2	.376280
29.75	.365257	29.5	.373009	29.45	.377054

Forster, 1878

%	t	n_D
0	24.0	1.33256
12.04	24.0	1.3496
23.61	24.8	1.36908

Wegner, 1889

%	n H_α	D	H_β	H_γ
		20°		
31.0222	1.381329	1.381825	1.390233	1.394904
20.4406	.361907	.362311	.365853	.373712
11.9564	.348152	.348488	.355299	.359215
4.9508	-	.338092	-	-
2.6741	-	.334901	-	-

Bender, 1890

M	n H_α	t	D	t
1	1.3440	21.3	1.34635	21.3
2	.35740	20.6	.35933	21.3
3	.36984	20.8	.37205	21.5
4	.38186	20.9	.38434	21.8
	H_β	t	H_γ	t
1	1.35103	21.3	1.35468	21.2
2	.36448	28.8	.36868	20.9
3	.37771	21.0	.38210	21.4
4	.39067	21.3	.39527	21.3

Jahn, 1891

%	n H_α	D	D'	H_β
		20°		
10.29	1.3461	1.3482	1.3481	1.3528
16.04	1.3551	1.3572	1.3573	1.3623

Borgesius, 1895

%	t	Dn	%	t	Dn
		20°			
1.67	18.4	0.000237	11.30	21.7	0.001744
3.29	18.7	.000472	11.30	18.0	.001751
6.02	16.9	.000889	12.09	20.7	.001883
6.33	19.6	.000926	20.15	16.5	.003524

Dn= n aq. - n sol (5890Å)

Le Blanc and Rohland, 1896

%	n_D
20°	
17.15	1.3590
22.72	1.3687

Wagner, 1903

c	n_D	c	n_D
17.5°			
0	1.33320	10.118	1.34687
0.278	.33358	10.401	.34721
0.556	.33397	10.684	.34761
0.834	.33435	10.967	.34798
1.112	.33474	11.251	.34836
1.390	.33513	11.524	.34873
1.670	.33551	11.817	.34910
1.950	.33590	12.100	.34947
2.230	.33628	12.385	.34984
2.510	.33667	12.666	.35021
2.790	.33705	12.950	.35058
3.071	.33743	13.234	.35095
3.352	.33781	13.518	.35132
3.633	.33820	13.802	.35169
3.914	.33858	14.086	.35205
4.195	.33896	14.370	.35242
4.476	.33934	14.654	.35279
4.757	.33972	14.938	.35316
5.038	.34010	15.222	.35352
5.319	.34048	15.506	.35388
5.600	.34086	15.790	.35425
5.882	.34124	16.074	.35461
6.164	.34162	16.358	.35497
6.446	.34199	16.642	.35533
6.728	.34237	16.926	.35569
7.010	.34275	17.211	.35606
7.292	.34313	17.496	.35642
7.574	.34350	17.781	.35678
7.856	.34388	18.066	.35714
8.138	.34426	18.351	.35750
8.420	.34463	18.636	.35786
8.703	.34500	18.921	.35822
8.986	.34537	19.206	.35858
9.269	.34575	19.491	.35894
9.552	.34612	19.776	.35930
9.835	.34650	20.061	.35966

Chéneveau, 1907

%	n_D	%	n_D
22°			
21.83	1.3663	11.88	1.3498
18.05	1.3598	9.68	1.3462
14.01	1.3531	5.02	1.3395

Baxter, Boylston and al., 1911

%	n_D	%	n_D
25°			
0	1.33246	4.6241	1.33890
0.4857	.33311	6.5288	.34159
0.9110	.33370	7.2055	.34257
1.1162	.33402	7.3936	.34291
1.9817	.33522	8.8063	.34501
2.1044	.33535	17.261	.35822
2.8104	.33639	46.06	.41841
2.9218	.33649		
4.5543	.33883		

Heydweiller, 1913

N	n_D
18°	
0.5	1.34031
1.0	.34718
2.0	.36021
4.0	.38611

Grufki, 1913

M	n H_α	n H_β	n H_γ
18°			
0	1.33144	1.33737	1.34058
0.5023	.33842	.34469	.34808
1.008	.34527	.35190	.35555
1.994	.35821	.36546	.36956
3.975	.38327	.39187	.39677

Kohner, 1928

N	n_D	N	n_D
25°			
0	1.33233	3.99924	1.37969
0.99522	.34574	4.99933	.38941
2.01264	.35807	6.02882	.39846
3.00158	.36516		

Shibata and Hölemann, 1931

%	$n_{He\ y}$	%	$n_{He\ y}$	%	$n_{He\ y}$
25.0°		35.0°		45.0°	
0	1.33270	0	1.33149	0	1.33000
12.13	.34786	15.56	.35088	14.13	.34747
29.92	.36848	26.02	.36269	25.66	.36048
42.85	.38160	53.46	.38965	53.48	.38768
53.48	.39160				

Guillaume, 1946

%	n_{5780}
20°	
9.46	1.3477
20.66	1.3661

Johnston, 1907

M	ϰ
100°	
4	5776
2	3622
1	2091
0.5	1213
0.25	683.5
0.1	287.3
0.05	151.3
0.0003	1.258

Heydweiller, 1909

M	λ
18°	
0.05	99.1
0.1	96.0
0.2	91.2
0.5	84.6
1.0	78.1
2.0	69.1
4.0	53.0

Gropp, 1915

t	ϰ
3.885 N	
0	1380
18	2086
48	3446
78	4912
100	5943

Humburg, 1893

%	(α) magn.
16°	
0	1
29.705	9.1852

Oppenheimer, 1898

%	rotatory magnetic specific polarization
20°	
9.41	1.62
17.77	1.62
25.10	1.66
33.57	1.62

Jahn, 1891

%	(α) magn.
20°	
124.74	16.04
115.17	10.29
99.88	0

Guillaume, 1946

%	* α magn. 10^6 (5780A°)
20°	
0	3.974
9.46	4.219
20.66	4.502

* in radians, gauss, centim.

Okazaki, 1933

%	Verdet's const.*	%	Verdet's const.
28°			
6.83	0.05004	22.05	0.06416
11.33	0.05416	29.70	0.07305
17.05	0.05939	36.26	0.08141

* Verdet's constant (λ = 3441 Å)

Okazaki, 1936

%	Verdet's const.*	%	Verdet's const.
	25°		
2.132	0.01349	17.51	0.01664
4.277	0.01389	21.90	0.01768
5.28	0.01404	22.51	0.01784
7.19	0.01446	28.98	0.01954
9.34	0.01489	33.81	0.02093
10.52	0.01511	38.44	0.02241
11.80	0.01538	40.15	0.02293

* Verdet's constant (D = 5890 Å)

Heat constants .

Faasch, 1911

N	U
18°	
0.673	0.934
1.347	0.873
2.69	0.777
5.39	0.649

Jauch, 1921

M	U
18°	
0.5	0.9448
1	0.8962
2	0.8156
3	0.7537
4	0.7054

Randall and Rossini, 1929

m	U	m	U
	25°		
0.00	0.9979	0.20	0.9741
0.01	.9966	0.35	.9576
0.02	.9954	0.50	.9420
0.05	.9917	0.75	.9177
0.10	.9857	1.00	.8953

Bender and Kaiser Jr., 1954

m	U	m	U
	25°		
0.9675	0.8983	3.8495	0.7247
0.9990	0.8955	5.9985	0.6534
2.2050	0.8078	8.3735	0.6028
2.2165	0.8076		
	30°		
0.9990	0.8973	5.9985	0.6528
2.2050	0.8102	8.3735	0.6020
3.8495	0.7256		

Jäger, 1891

%	τ (heat conductivity)
0	100
20	93
40	88.9

Kapustinskii and Ruzavin, 1956

m	H	m	H
	25°		
0.64	0.106	3.42	0.102
1.00	0.107	5.76	0.098
2.10	0.106		

$$H = \frac{1000 + mM/c - 1000\ co}{m}$$

m = molality

c = heat conductivity coefficient of the solution

co = " " " of water

Kapustinskii and Ruzavin, 1955

%	$c.10^6$	%	$c.10^6$
	25°		
4.16	1436	21.8	1355
5.30	1429	25.9	1334
9.33	1419	26.2	1332
10.8	1409	37.3	1264
13.5	1397	37.5	1266
17.6	1376		

c = heat conductivity coefficient

Wüst and Lange, 1925

mol%	Q diss integral	mol%	Q diss intermediate
	25°		
1.827	+209.2	0.913	+209.2
2.775	316.9	2.292	516.8
4.508	491.2	3.641	757.4
6.222	632.7	5.355	980.3
7.869	741.5	7.041	1127
9.471	822.8	8.678	1175
10.882	867.1	10.180	1131
13.177	884.7	11.626	996.4
3.010	+352.4	1.537	+352.9
4.980	541.5	4.013	815.3
6.888	687.7	5.946	1046
8.630	788.3	7.769	1153
10.291	854.5	9.463	1160
11.902	886.0	11.102	1065
13.458	885.4	12.693	996.4

Water (H_2O) + Sodium iodide (NaI)

Heterogeneous equilibria .

Tammann, 1885

t	p 0%	11.68%	30.54%	36.50%	46.48%
31.20	34.1	-	-	28.7	24.9
33.81	39.5	-	-	33.2	28.7
38.29	50.5	48.6	44.5	42.2	36.7
45.11	72.3	69.7	63.9	60.3	52.3
48.13	84.3	81.5	74.5	70.7	61.0
49.88	92.0	89.5	81.7	77.7	67.0
53.34	108.9	105.4	96.2	91.3	88.0
56.29	125.6	-	110.6	105.2	91.3
58.32	138.2	134.0	122.1	115.9	100.3
60.82	155.2	150.2	137.9	129.9	112.5
64.23	181.3	176.2	160.0	151.8	131.6
68.32	217.3	211.0	191.4	181.6	158.3
70.62	240.2	233.5	212.0	201.0	174.1
72.94	265.2	258.2	234.1	222.3	192.2
75.60	296.7	288.2	261.3	247.5	214.3
80.33	360.2	360.3	318.6	300.9	261.1
82.44	392.1	381.8	346.1	327.8	284.0
84.83	430.9	419.7	479.5	459.7	312.3
87.11	470.9	458.0	415.0	392.7	341.3
90.16	529.2	515.7	467.2	442.7	384.9
92.70	582.3	566.2	513.1	485.7	423.8
95.41	643.6	626.6	567.7	537.7	468.4
97.46	693.5	673.1	609.9	578.7	503.0
98.92	731.2	-	642.0	609.1	529.3
100.42	771.5	748.1	677.2	643.0	560.8

%	p	%	p
	100°		
51.24	497.8	68.40	263.4
56.69	428.1	70.98	223.4
67.14	281.8		

Lannung, 1934

m	p	m	p
	18°		
2.177	14.2	5.154	11.9
2.424	14.0	8.89	8.5
2.870	13.2	10.42	7.1
3.482	12.5	sat. sol.	6.1
(2 + 1)	tr.t. = 1.6		

Keevil, 1942

mol%	t	P	mol%	t	P
		sat.sol.			
33.6	185.0	2.40	55.9	389.5	38.79
39.0	240.8	8.51	62.4	444.0	61.2
45.4	300.2	17.58	69.0	493.0	88.5
50.2	342	26.36	73.5	521.4	100.04
53.8	372	34.63	88.8	600.0	189.1

Lescoeur, 1894

t	dissociation p. NaI.4aq.	p sat.sol.
20	-	5.4
30	4.3	-
40	9	17.3
50	17	20
60	28	28
70	-	45
80	-	70
90	-	104
100	-	150
110	-	240

Dingemans, 1938

t	p NaI.2aq. +NaI.+V	NaI.2aq. +sat.sol.+V	NaI.+ sat.sol.+V	L + V aq.
10.0	0.8	3.8	-	9.2
12.00	-	4.2	-	10.5
15.00	-	5.1	-	12.8
18.00	1.6	6.0	-	15.5
20.00	1.9	6.7	-	17.5
23.00	-	7.9	-	21.1
25.00	-	8.8	-	23.8
27.00	-	9.7	-	26.7
30.00	4.0	11.4	-	31.8
32.00	-	12.6	-	35.7
35.00	-	14.5	-	42.2
37.00	-	16.0	-	47.1
40.00	8.2	18.2	-	55.3
42.00	-	19.9	-	61.5
45.00	11.6	22.5	-	71.9
47.00	-	24.5	-	79.6
50.00	15.7	27.5	-	92.5
52.00	17.9	29.6	-	102.1
55.00	21.5	32.9	-	118.0
57.00	24.5	35.1	-	129.8
60.00	29.5	38.4	-	149.4
62.00	33.1	40.5	-	163.8
65.00	39.6	43.8	-	187.5
67.00	44.5	46.1	-	205.0
68.08	-	-	-	214.4
70.00	-	-	51.6	233.7
72.00	-	-	56.7	254.6
75.00	-	-	64.4	289.1
77.00	-	-	70.4	314.1
80.00	-	-	79.8	355.1
82.00	-	-	87.0	384.9
85.00	-	-	98.6	433.6
87.00	-	-	107.1	468.7
90.00	-	-	120.8	525.8
92.00	-	-	130.6	567.0
95.00	-	-	146.7	633.9
97.00	-	-	159.7	682.1
100.00	-	-	178.7	760.0

Schlamp, 1894

%	b.t.	%	b.t.
0	100	13.92	101.128
2.89	100.19	16.67	101.436
5.99	100.412	19.22	101.750
10.40	100.780		

Johnston, 1906

%	b.t.	%	b.t.
4.29	100.281	50.73	111.610
13.39	101.082	53.32	113.334
24.03	102.427	54.71	114.359
29.23	103.412	57.64	116.244
35.30	104.794	58.75	117.224
41.39	106.696	59.54	118.172
45.41	108.314	60.32	118.594
48.29	110.014		

Jablezynski and Kon, 1923

m	b.t.	m	b.t.
0.3432	100.334	1.3889	101.509
0.6927	100.702	1.5819	101.752
1.0038	101.051	1.7691	101.997
1.2957	101.286	1.9644	102.260

Briggs and Geigle, 1940

%	b.t.	%	b.t.
37.2	104.8	64.0	121.6
44.2	107.0	70.0	131.8
53.3	113.0	75.0	142.0
57.2	117.4	76.2	142.4

Robinson, 1935

Isopiestic solutions	m_k	m_i	m_k	m_i
		25°		
	0.1158	0.1146	1.622	1.431
	0.2014	0.1969	1.938	1.680
	0.2029	0.1986	2.214	1.898
	0.2708	0.2616	2.313	1.974
	0.4010	0.3834	2.565	2.165
	0.4031	0.3858	2.909	2.425
	0.5590	0.5271	2.985	2.479
	0.6880	0.6409	3.096	2.573
	0.7132	0.6650	3.229	2.655
	0.7615	0.7078	3.344	2.738
	0.8900	0.8199	3.569	2.905
	0.8944	0.8208	3.823	3.074
	1.078	0.9826	4.139	3.292
	1.112	1.010	4.290	3.388
	1.245	1.120	4.564	3.577
	1.314	1.176	4.81	3.740

m_k = molality of potassium chloride

m_i = molality of sodium iodide

Miller and Sheridan, 1956

Isopiestic solutions	m_i	m_s	m_i	m_s
		25°		
	3.94	2.32	9.29	6.89
	7.10	5.16	11.90	7.96
	9.19	6.54	12.72	8.52

m_i = molality of sodium iodide .
m_s = molality of sulfuric acid .

Robinson and Stokes, 1956

m	osmotic coefficient	m	osmotic coefficient	m	osmotic coefficient
		25°			
0.1	0.938	0.7	0.967	1.6	1.043
0.2	0.936	0.8	0.975	1.8	1.061
0.3	0.939	0.9	0.983	2.0	1.079
0.4	0.945	1.0	0.991	2.5	1.129
0.5	0.952	1.2	1.007	3.0	1.188
0.6	0.959	1.4	1.025	3.5	1.243

Kremers, 1856

%	f.t.	%	f.t.	%	f.t.
61.35	0	71.94	60	76.34	120
64.10	20	75.19	80	76.92	140
67.57	40	75.76	100	sat.sol.	141

Rüdorff, 1862

%	f.t.	%	f.t.
3.61	-0.85	14.59	-4.25
7.48	-1.9	26.50	-9.75

Guthrie, 1876

%	f.t.	%	f.t.	%	f.t.
5	- 0.7	30	-11.8	50	-29.5
10	- 2.1	35	-15.2	55	-20.0
15	- 3.9	40	-20.5	60	-14.7
20	- 6.0	45	-26.0	61.6	0.0
25	- 8.5	49.2	-30.0 E	63.6	+13

De Coppet, 1883

%	f.t.	%	f.t.	%	f.t.
59.89	-17.25	63.56	+14.1	75.02	97.1
60.81	- 5.4	74.65	64.7	75.15	101.7
61.43	0.0	74.64	71.3	75.37	110.7
61.89	+ 3.15	74.70	74.1	76.03	124:7
62.08	+ 4.95	74.79	81.6	76.02	132.5
63.46	+12.5	74.90	86.4	76.15	138.1
62.47	+12.6	75.01	92.4		

Jones and Getman, 1904

M	f.t.	M	f.t.
0.083	-0.314	2.007	- 8.700
0.167	-0.595	2.676	-12.720
0.334	-1.186	3.345	-18.000
0.669	-2.437	4.014	-23.000
1.338	-5.370	4.683	-29.500

Meyerhoffer, 1905

39 % f.t. = -31.5° NaI. 5 aq.

Scott and Frazier, 1927

64.76 % (sat. sol.) f.t. = 25°

Scott and Durham, 1930

%	f.t.	%	f.t.
61.54	0.00	74.76	67.93
66.35	35.00	75.02	70.87
69.42	50.02	74.82	75.40
72.95	63.13	75.05	92.23

Briggs and Geigle, 1940

%	f.t.	E	%	f.t.	E
18.7	-5.6	-	53.3	-24.8	-31.5
37.2	-17.8	-	57.2	-17.3	-31.8
41.9	-23.5	-31.7	61.0	-12.2	-31.8
44.2	-26.8	-31.2	64.0	-12.4	-
46.5	-30.6	-31.5	77.5	68.0	-
49.0	-36.1	-31.6	82.4	68.0	-

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
NaI.2aq.		NaI.x aq.	
16.29	2.3	24.81	46.16
17.14	13.08	25.23	51.64
19.63	37.67	25.79	58.30
21.19	47.97	26.08	72.29
23.37	58.93	26.32	85.2
24.35	62.78	26.61	103.0
24.81	64.57	27.28	126.9
25.23	65.68	27.94	147.7
25.79	67.63	29.31	182.9
26.08	68.43		

Properties of phases.

Kremers, 1855, 1858 and 1859

%	d	%	d
	19.5°		
5	1.038	35	1.358
10	.080	40	.430
15	.126	45	.527
20	.177	50	.597
25	.232	55	.697
30	.292	60	.807

t	d			
0	1.2303	1.4269	1.6086	1.7766
19.5	.2213	.4139	.5915	.7574
40	.2099	.3977	.5731	.7369
60	.1971	.3818	.5547	.7165
80	.1828	.3651	.5360	.6960
100	.1675	.3479	.5169	.6759

% from former table .

Kohbrausch, 1879

%	d	%	d
	18°		
5	1.0374	30	1.2836
10	.0803	40	1.4127
20	.1735		

Röntgen and Schneider, 1886

%	d
18.1°	
18.75	1.1647
9.51	1.0781

Jahn, 1891

%	d
20°	
21.78	1.1929
14.30	1.1190

Jones and Getman, 1904

%	d	%	d
		0°	
1.236	1.007316	24.58	1.224964
2.464	1.016508	30.82	1.302584
4.835	1.036228	36.47	1.375704
9.339	1.064564	41.56	1.448736
17.43	1.151424	46.01	1.526748

Agerer, 1905

%	d
18°	
6.1190	1.0460
13.382	.1102
25.861	.2379
42.247	.4558

Chéneveau, 1907

%	d	%	d
		22°	
22.06	1.1960	9.79	1.0794
18.25	1.1579	5.08	1.0400
14.16	1.1182	0	0.9978
12.12	1.0986		

Heydweiller, 1909

M	d	M	d
		18°	
0.05	1.00431	1.0	1.1108
0.1	.0100	2.0	.2209
0.2	.0212	4.0	.4373
0.5	.0549		

Grufki, 1913

M	d	M	d
		18°	
0	0.99862	2.261	1.25387
0.5068	1.05630	4.301	1.47839
1.015	1.11438		

Lübben, 1914

M	d	M	d
		18°	
0.626	1.06885	2.476	1.28183
1.248	.14025	3.740	.42754

Gropp, 1915

t	d	t	d
	6.789 %		
0	1.7573	78	1.6870
18	.7417	100	.6660
48	.7148	108	.6545

Baxter and Wallace, 1916

%	d	%	d
70.19°		25°	
5.05	1.01599	58.75	1.76970
50.04°		37.42	.38569
		18.00	.15413
		9.60	.07550
59.20	1.75460	4.96	.03624
37.75	.37192	0°	
18.17	.14261		
9.68	.06519		
5.00	.02650	58.30	1.78425
		36.93	.39837
		17.88	.16237
		9.56	.08136
		4.94	.04067

Rabinowitsch, 1921

%	d	%	d
	25°		
58.08	1.7545	22.20	1.1970
50.39	.5965	14.45	.1195
40.41	.4297	7.75	.0592
34.80	.3403	2.94	.0196
30.07	.2880		

de Block, 1925

%	d	%	d
	15°		
0	0.9991	35.5	1.3652
15.0	1.1151	50.5	.6069
20.4	1.1802	62.4	.8761

Hrynakowski, 1927

sat.sol. (50°) d= 2.0166

Scott and Frazier, 1927

64.76% 25° d=1.91901

Scott and Durham, 1930

%	t	d	%	t	d
		sat.sol.			
61.54	0.00	1.8612	74.76	67.93	2.1554
66.35	35.00	1.9512	75.02	70.87	.1577
69.42	50.02	2.0169	74.82	75.40	.1544
72.95	63.13	2.1068	75.05	92.23	.1425

Geffcken, 1929

m	d	m	d
	25°		
0	0.99707	6.1257	1.55153
1.5655	1.16436	9.0524	1.74315
3.4363	1.33883		

Scott, Obenhaus and Wilson, 1934

%	d
35°	
60.802	1.80644
40.413	.42173
22.712	.19758
13.896	.10998

Freyer, 1931

%	π	%	π
	20°		
6	44.05	30	35.18
10	42.64	40	31.21
20	38.97	45	28.97

t	π relative	t	π relative
	45%		
10	28.77	30	29.09
20	28.97	40	29.26

Gibson, 1937

%	π (1-1000 kg.)	%	π (1-1000 kg.)
0.000	39.35	44.546	27.65
10.016	37.01	49.706	26.05
19.740	34.61	50.483	25.69
34.424	30.61		

Scott, Obenhaus and Wilson, 1934

%	π
35°	
60.802	24.22
40.413	31.39
22.712	36.36
13.896	38.55

Röntgen and Schneider, 1886

%	π relative
18.0°	
18.75	0.859
9.51	0.924

Dunlop and Stokes, 1951

M	D (cm^2 sec^{-1})	M	D (cm^2 sec^{-1})
25°			
0.04193	1.554	0.6142	1.560
.04396	.549	.6237	.559
.03608	.535	.9324	.579
.03311	.539	.9324	.571
.08916	.535	1.412	.614
.09011	.538	.425	.612
.09073	.535	.675	.641
.09181	.537	.982	.668
.1729	.529	.998	.667
.1759	.533	2.341	.701
.2697	.534	.364	.701
.2740	.534	.801	.742
.3743	.534	.835	.739
.3874	.539	3.439	.785
.4434	.541		

D = integral diffusion coefficient

Hechler, 1904

t	η	t	η
59.5 %			
-19.61	8174	-8.20	4387
19.09	7676	6.21	4040
18.07	7038	+0.06	3130
17.03	6712	3.11	2810
16.06	6400	5.72	2550
-14.54	5932	+7.90	2420

Rabinowitsch, 1921

%	η	%	η
25°			
58.08	2420	22.20	1067
50.39	1728	14.45	1026
40.41	1324	7.75	1014
34.80	1191	2.94	1005
30.07	1142	0	1000

Dunlop and Stokes, 1951

M	η (water =1)	M	η
25°			
0.290	1.008	2.059	1.094
0.607	1.015	2.947	1.201
1.050	1.032		

Hrynakowski, 1927

sat.sol. (50°) η = 2669 σ = 80.42

de Block, 1925

%	σ	%	σ
15°			
0	73.26	35.5	74.88
15.0	73.58	50.5	77.50
20.4	73.79	62.4	81.44

Kremers, 1857

%	t	n_D
42.17	17	1.4167
58.04	18	1.4786

Bender, 1890

M (15°)	n H_α	t	n D	t
1	1.35182	19.6	1.35390	20.2
2	.37204	19.9	.37452	19.1
3	.39196	19.6	.39484	20.0
4	.41184	19.8	.41501	21.1
5.3865	.43848	20.6	.44243	21.4
	H_β	t	H_γ	t
1	1.35925	19.5	1.36326	20.0
2	.38082	20.0	.38593	20.0
3	.40242	19.2	.40836	19.6
4	.42358	20.0	.43062	20.0
5.3865	.45229	20.9	.46063	21.0

Jahn, 1891

%	n H_α	n D	n D'	n H_β
		20°		
21.78	1.3675	1.3700	1.3700	1.3760
14.30	1.3536	1.3559	1.3559	1.3612

Chéneveau, 1907

%	n_D	%	n_D
	22°		
22.06	1.3690	9.79	1.3470
18.25	.3617	5.08	1.3398
14.16	.3544	0	1.3328
12.12	.3508		

Wagner, 1903

c	n_D	c	n_D
	17.5°		
0	1.33320	10.541	1.34798
0.267	.33358	10.801	.34836
.534	.33397	11.061	.34873
.801	.33435	11.321	.34910
1.067	.33474	11.581	.34947
.333	.33513	11.840	.34984
.599	.33551	12.099	.35021
.865	.33590	12.358	.35058
2.131	.33628	12.617	.35095
.396	.33667	12.876	.35132
.661	.33705	13.135	.35169
.926	.33743	13.393	.35208
3.191	.33781	13.651	.35242
3.456	.33820	13.909	.35279
3.721	.33858	14.167	.35316
3.985	.33896	14.425	.35352
4.249	.33934	14.683	.35388
4.513	.33972	14.940	.35425
4.777	.34010	15.197	.35461
5.041	.34048	15.454	.35497
5.305	.34086	15.711	.35533
5.568	.34124	15.968	.35569
5.831	.34162	16.225	.35606
6.094	.34199	16.481	.35642
6.357	.34237	16.737	.35678
6.620	.34275	16.993	.35714
6.883	.34313	17.249	.35750
7.145	.34350	17.505	.35786
7.407	.34388	17.761	.35822
7.669	.34426	18.016	.35858
7.931	.34463	18.271	.35895
8.193	.34500	18.526	.35930
8.716	.34537	18.781	.35966
8.977	.34575	19.036	.36002
9.238	.34612	19.291	.36038
9.499	.34650	19.545	.36074
9.760	.34687	19.799	.36109
10.021	.34724	20.053	.36145
10.281	.34761	20.307	.36181

Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.083	1.32735	2.676	1.38319
.167	.32925	3.345	.39639
.334	.33280	4.014	.41053
.669	.34004	4.683	.42481
1.338	.35481	5.352	.43903
2.007	.36868	6.693	.46613

Baxter, Boylston and al., 1911

%	n_D	%	n_D
	25°		
0.5723	1.33333	3.8603	1.33817
0.7125	.33352	6.8272	.34278
0.7137	.33353	18.4384	.36281
0.7469	.33362	30.418	.38783
3.7111	.33793		

Grufki, 1913

M	n H_α	H_β	H_γ
		18°	
0	1.33140	1.33735	1.34054
0.5068	.34210	.34878	.35248
1.015	.35288	.36036	.36456
2.261	.37842	.38773	.39316
4.301	.41915	.43152	.43879

Heydweiller, 1913

M	n_D
18°	
0.5	1.34411
1.0	1.35476
2.0	1.37585
4.0	1.41763

Lübben, 1914

M	n 18° 4600.9 Å	4416.2 Å	3848.9 Å	3749.6 Å
0	1.35334	1.35474	1.34477	1.34586
0.626	.36769	.36846	.37612	.36203
1.248	.39537	.39749	.37630	.37788
2.476	.42317	.42581	.40661	.40863
3.740	.33878	.33992	.43710	.43956
	3611.9 Å	3467.0 Å	3403.6 Å	3333.7 Å
0	1.34755	1.34961	1.35061	1.35179
0.626	.36414	.36671	.36798	.36950
1.248	.38042	.38353	.38504	.38690
2.476	.41199	.41612	.42816	.42067
3.740	.44380	.44895	.45153	.45473
	3255.8 Å	3133.3 Å	3085.1 Å	2981.1 Å
0	1.35322	1.35572	1.35681	1.35944
0.626	.37133	.37463	.37605	.37959
1.248	.38911	.39318	.39490	.39939
2.476	.42359	.42934	.43165	.43797
3.740	.45854	.46579	.46874	.47771
	2881.1 Å	2837.1 Å	2748.7 Å	
0	1.36224	1.36358	1.36653	
0.626	.38351	.38544	.38990	
1.248	.40437	.40686	.41274	
2.476	.44509	.44855	.45742	
3.740	.48610	.49072	.50228	

Geffcken, 1929

m	n_D	m	n_D
	25°		
0	1.332590	6.1257	1.365772
1.5655	.363890	9.0524	.470916
3.4363	.366317		

Kohlrausch, 1879

%	κ	$\tau.10^4$
	18°	
5	296	222
10	577	216
20	1136	204
30	1643	198
40	2097	198

Hechler, 1904

t	κ	t	κ
	59.5%		
-19.62	607	8.28	969
19.15	626	6.30	1040
18.15	651	0.00	1287
17.11	681	+2.42	1385
16.11	707	3.09	1385
14.42	763	5.65	1523
13.33	797	7.87	1620
11.87	844		

Jones and Getman, 1904

M	λ	M	λ
	0°		
0.083	58.66	2.676	41.77
0.167	56.75	3.345	37.41
0.334	53.21	4.014	34.97
0.669	50.28	4.683	30.83
1.338	46.50	5.352	26.43
2.007	45.50	6.693	19.69

Johnston, 1906

M	λ	M	λ
	99.4°		
8.33	855	0.5	2206
7	953	0.25	2524
5	1307	0.10	2811
3	1745	0.001	3728
1	-		

Heydweiller, 1909

M	λ	M	λ
	18°		
0.05	97.7	1.0	78.6
0.1	94.0	2.0	70.2
0.2	90.2	4.0	53.9
0.5	84.1		

Gropp, 1915

t	κ	t	κ
	7 N		
-25.0	494.7	0	1286
-20.5	614.3	18	2112
-12.4	852.4	48	3735
-9.5	928.8	78	5315
		100	6641

Rabinowitsch, 1921

%	κ	%	κ
	25°		
58.08	2504	22.20	1484
50.39	2671	14.45	978
40.41	2475	7.75	532
34.80	2193	2.94	209
30.07	1967	0	0

Miller, 1956 (fig.)

M	λ		
	0°	29.9°	50°
1	52	-	135
2	47	89	120
4	35	68	92
8	14	31	42

Jahn, 1891

%	(α) magn.
21.78	150.80
14.30	131.98

Agerer, 1905

%	t	(α) magn.
6.1190	18.3	2.320
13.382	19.5	.308
25.861	19.0	.282
42.247	19.4	.259

Okazaki, 1933

%	Verdet's constant (3441 Å)
28°	
6.94	0.05378
14.19	.06573
20.70	.07691
28.25	.09225
34.53	.10626

Okazaki, 1936

%	Verdet's constant (5780Å)
25°	
0	0.013075
7.015	.01500
14.07	.01727
22.18	.02012
28.06	.02248
35.38	.02578
43.14	.02988
52.27	.03562
54.57	.03729
59.57	.04107

Heat constants.

Schüller, 1869

%	U	%	U
28.57	0.7343	16.67	0.8408
23.08	0.7811	9.09	0.9135

Faasch, 1911

N	U	N	U
	18°		
0.608	0.902	2.450	0.724
1.218	.840	4.94	0.575

Jauch, 1921

N	U	N	U
	18°		
0.5	0.9231	3	0.6769
1	.8604	4	.6156
2	.7554		

Randall and Rossini, 1929

N	U	N	U
	25°		
0.00	0.9979	0.20	0.9654
0.01	.9962	0.35	0.9430
0.02	.9945	0.50	0.9220
0.05	.9894	0.75	0.8897
0.10	.9812	1.00	0.8603

Wüst and Lange, 1925

mol%	Q diss. (integral)	mol%	Q diss. (integral)
	25°		
1.396	+1868	0.698	+1868
2.770	2037	2.070	2206
4.060	2192	3.405	2505
5.330	2329	4.718	2740
6.707	2446	5.948	2923
7.729	2543	7.147	3041
8.911	2622	8.838	3089
10.059	2684	9.484	3125
11.182	2730	10.639	3079
12.274	2758	11.729	3002
13.332	2766	12.812	2856
14.390	2757	14.540	2662
15.043	2743	14.715	2489

Kapustinskii and Ruzavin, 1955

%	$c.10^6$	%	$c.10^6$
	25°		
5.10	1435	27.8	1298
10.9	1403	42.9	1199
10.7	1365	49.9	1137

c= Heat conductivity coefficient

Kapustinskii and Ruzavin, 1956.

m	H	m	H
	25°		
0.2	0.16	2.5	0.135
0.8	0.15	5.0	0.128
1.3	0.145	6.6	0.123

$$H = \frac{(1000 - mM)\,c - 1000\,c_0}{m}$$

m = molality

c = heat conductivity coefficient of the solution

c_0 = " " " of water

Water (H_2O) + Sodium Azide (NaN_3)

Wohlgemuth, 1934

%	f.t.	
0	0	
10	-7	
21.6	-15.1 E	ice + NaN_3 . 3 aq.
26.8	-20.0 E	metast. ice + NaN_3
27.8	-2.1	
30	+40	
35	100	

Günther and Perschke, 1930

M	d	η	n_D
	20°	20°	25°
3.0	1.1137	744.0	1.38319
2.0	.0764	848.1	.36699
1.0	.0380	952.0	.35002
0.5	.0184	994.0	.34131

Water (H_2O) + Sodium thiocyanate (NaCNS)

Tammann, 1915

%	p	%	p
	100°		
4.62	746.6	33.73	554.5
9.41	726.8	42.54	470.8
10.16	724.4	47.71	410.6
27.56	614.9	48.92	398.6
28.58	605.1	50.51	380.0

Miller and Sheridan, 1956

Isopiestic solutions

m_1	m_2	m_1	m_2
	25°		
4.40	3.06	12.44	7.80
6.40	4.51	14.11	8.50
7.22	4.80	15.20	9.12
8.36	5.77	17.86	10.15
10.30	6.86	19.01	10.85

m_1 = molality of sodium thiocyanate

m_2 = molality of sulfuric acid

Robinson and Sotkes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.937	1.2	0.979
0.2	0.934	1.4	0.990
0.3	0.935	1.6	1.002
0.4	0.938	1.8	1.014
0.5	0.943	2.0	1.025
0.6	0.948	2.5	1.055
0.7	0.953	3.0	1.086
0.8	0.958	3.5	1.118
0.9	0.963	4.0	1.150
1.0	0.969		

Grufki, 1913

M	%	d
	18°	
0	0	0.99862
0.5054	3.52	1.01968
0.829	5.69	.03303
1.946	12.82	.07864
3.990	24.56	.15548

Hughes and Mead, 1929

%	f.t.	%	f.t.
52.5	10.7	63.82	46.1
55.62	17.3	65.00	65.8
58.00	21.3	66.21	73.8
62.35	29.2	66.50	81.8
63.20	33.8	69.25	101.4

Bump, 1932

%	f.t.
61.75	15
62.31	20
62.45	25

Dixon and Taylor, 1910

mol %	d
18°	
0	0.9986
8.15	1.1600
13.04	1.2306
15.09	1.2599

Heydweiller, 1921

M	d
18°	
0.5	1.02087
1	1.0414
2	1.0813
3	1.1197
4	1.1573

Miller and Doran, 1956

0° : $d = 0.9999 + 0.04823\ M - 0.003472\ M^{3/2}$

29.87° : $d = 0.9957 + 0.04137\ M - 0.001600\ M^{3/2}$

50.1° : $d = 0.9880 + 0.03778\ M - 0.0002084\ M^2$

M = moles per l. at t.

m	η (water=1) 0°	29.87°	50.1°
0	1.2	1.2	1.2
2	1.25	1.3	1.3
4	0.7	0.75	0.75
6	2.6	2.5	2.5
8	5.1	4.2	4.0
9	9.5	6.7	5.8
10	-	9.0	7.7

Dixon and Taylor, 1910

mol %	n_D
18°	
0	1.33314
8.15	1.40250
13.06	1.43336
15.09	1.44461

Heydweiller, 1913

M	n_D
18°	
0	1.33327
0.5	.33990
1.0	.34642
2.0	.35927
4.0	.38387

Grufki, 1913

M	%	n Hα	n Hβ	n Hγ
		18°		
0	0	1.33143	1.33740	1.34055
0.5054	3.52	.34024	.34668	.35017
0.829	5.69	.34579	.35246	.35616
1.946	12.82	.36469	.37239	.37674
3.990	24.56	.39779	.40732	.41279

Heydweiller, 1921

M	λ
18°	
0.5	74.3
1	68.9
2	59.8
3	50.8
4	43.70

Miller, 1956 (fig.)

M	λ 0°	λ 29.9°	λ 50°
0	53	112	152
2	38	75	125
4	26	57	78
8	11	20	33
9.5	6	16	23

Partington and Soper, 1929

mol%	Q diss(by mole NaCNS) integral	Q diss(by mole NaCNS) differential	Q dil(by mole H_2O) differential
0	-1677	-1677	0.0
0.4	1648	1571	-0.38
0.9	1574	1426	1.48
1.4	1500	1268	3.48
1.9	1425	1113	6.24
16.6	434	-	-
17.0	439	-	-
17.3	444	-	-
17.8	450	-	-
18.0	455	612	+36.30
18.3	460	625	37.12
18.7	465	640	40.25
19.1	470	651	42.50
19.3	476	687	50.64
19.6	481	716	57.57
20.0	486	751	66.25
20.4	493	788	75.225
20.6	500	827	85.02
20.9	507	868	95.665
21.2	514	909	106.65
21.6	521	947	117.02

Water (H_2O) + Sodium sulfide (Na_2S)

Parravano and Fornaini, 1907

f.t.	%	f.t.	%
	9 aq.		
10	13.36	28	17.73
15	14.36	32	19.09
18	15.30	37	20.98
22	16.20	45	24.19
	6 aq.		
50	26.70	80	32.95
60	28.10	90	36.42
70	30.22		
	5.5 aq.		
50	28.48	70	31.38
55	29.27	80	33.95
60	29.92	90	37.20

Sanfourche and Liebaut, 1922

%	f.t.	tr.t.	E	
1.8	-2	-	-9.5	
4.7	-3.6	-	"	
7.5	-9.5	-	"	
9.5	+9	-	"	
14.0	17	-	"	
21.0	40	-	-8	
23.0	44	-	"	
26.0	47	-	-9	
28.0	50	50	-	
30.5	62	"	-9	
39.3	90	51	"	
43.6	97	50	"	(5+1)
44.7	98	-	"	
44.8	97.5	-	"	
45.3	97.5	-	"	
46.0	97	-	85	
50.4	93	-	"	
52.8	88	-	"	
53.5	"	-	"	
54.4	86	-	"	
54.6	"	-	"	
55.5	"	-	88	
56.2	87	-	85	
58.4	90.5	-	"	
60.1	95	-	"	
61.4	-	94	"	
63.2	-	"	"	
63.5	-	92	"	
64.7	-	93.5	"	
65.9	-	91	"	
67.9	-	93	"	
77.6	-	94	"	
79.6	-	"	-	

Grube, 1927

f.t.	%
14.10	14.99
18.10	16.10
22.00	17.26
26.10	18.55
30.05	19.80

Nasini and Costa, 1891

27.2976 % d = 1.17889

Grube, 1927

%	d		
	15°	18°	21°
14.99	1.15725	1.15566	1.15407
16.10	.16829	.16672	.16519
17.26	.18009	.17853	.17696
18.55	.19363	.19201	.19039
19.80	.20922	.20756	.20590

%	κ
	15°
14.99	2468
16.10	2464
17.26	2442
18.55	2385
19.80	2295

Nasini and Costa, 1891

%	n		
	Hα	Hβ	Hγ
	26.2°		
27.2976	1.38161	1.38896	1.39263

XLV. WATER + SODIUM MONOXYSALTS .

Water (H_2O) + Sodium hydroxide (NaOH)

Heterogeneous equilibria .

Wüllner, 1860

t	p			
	0%	9.09%	16.67%	23.08%
14.50	12.298	11.051	9.655	8.458
20.20	17.608	15.963	14.064	12.064
22.73	20.594	19.100	16.859	15.664
25.06	23.642	22.850	19.610	16.922
27.88	27.900	25.860	22.975	20.040
30.72	32.869	30.283	27.150	23.824
31.05	33.499	30.519	26.731	24.251
32.80	36.991	34.061	30.431	26.655
34.65	41.021	38.085	33.921	29.773
35.66	41.953	38.518	34.591	30.060
37.93	49.284	45.413	40.993	36.008
38.30	50.021	46.207	41.609	36.595
40.50	56.406	52.354	47.035	41.323
42.70	63.355	58.702	54.702	46.327
43.68	66.685	61.483	55.627	49.085
45.68	73.947	68.647	61.903	58.612
48.03	83.325	77.377	69.585	61.328
49.28	89.229	82.094	74.308	65.635
50.75	95.492	88.772	80.292	71.065
52.75	105.287	97.675	88.390	78.287
54.28	113.333	104.379	94.934	83.679
56.05	123.544	114.690	103.942	92.219
57.65	133.313	123.764	112.059	99.404
58.66	139.630	129.432	117.335	104.294
60.28	150.553	139.357	126.410	112.267
62.40	166.218	154.214	139.940	120.683
64.30	181.181	167.809	151.825	135.863
65.42	190.345	176.683	160.395	142.776
67.60	210.000	195.032	176.847	157.844
69.36	226.165	209.611	190.560	169.764
70.65	239.788	222.575	202.610	-
72.00	254.073	235.484	214.240	-
73.40	269.727	250.070	227.763	-
75.43	293.814	272.392	248.017	-
77.15	315.581	292.281	265.725	-
79.33	346.206	321.096	292.914	-
81.16	371.709	344.335	313.954	-
82.58	393.517	365.321	332.993	-
83.38	406.253	376.581	343.240	-
84.70	428.016	395.795	361.092	-
85.63	443.827	412.227	375.827	-
87.15	470.655	436.886	398.170	-
89.43	514.215	476.768	335.715	-
91.25	550.935	510.598	466.386	-
93.28	594.596	552.258	504.084	-
95.43	643.890	597.979	546.451	-
97.20	687.020	634.037	579.670	-
99.50	747,500	693.350	633.205	-

Tammann, 1885

%	p	%	p	%	p
		100°			
1.72	749.8	18.30	601.4	27.96	453.7
2.90	742.5	18.88	579.5	31.05	396.2
7.23	712.8	21.83	553.7	36.66	307.6
10.85	681.3	24.62	511.2	42.81	217.0
12.54	665.5				

Grube, 1927

f.t.	%
14.10	14.99
18.10	16.10
22.00	17.26
26.10	18.55
30.05	19.80

Nasini and Costa, 1891

27.2976 %   d = 1.17889

Grube, 1927

%	d		
	15°	18°	21°
14.99	1.15725	1.15566	1.15407
16.10	.16829	.16672	.16519
17.26	.18009	.17853	.17696
18.55	.19363	.19201	.19039
19.80	.20922	.20756	.20590

%	κ
	15°
14.99	2468
16.10	2464
17.26	2442
18.55	2385
19.80	2295

Nasini and Costa, 1891

%	n		
	Hα	Hβ	Hγ
	26.2°		
27.2976	1.38161	1.38896	1.39263

XLV. WATER + SODIUM MONOXYSALTS .

Water (H_2O) + Sodium hydroxide (NaOH)

Heterogeneous equilibria .

Wüllner, 1860

t	p			
	0%	9.09%	16.67%	23.08%
14.50	12.298	11.051	9.655	8.458
20.20	17.608	15.963	14.064	12.064
22.73	20.594	19.100	16.859	15.664
25.06	23.642	22.850	19.610	16.922
27.88	27.900	25.860	22.975	20.040
30.72	32.869	30.283	27.150	23.824
31.05	33.499	30.519	26.731	24.251
32.80	36.991	34.061	30.431	26.655
34.65	41.021	38.085	33.921	29.773
35.66	41.953	38.518	34.591	30.060
37.93	49.284	45.413	40.993	36.008
38.30	50.021	46.207	41.609	36.595
40.50	56.406	52.354	47.035	41.323
42.70	63.355	58.702	54.702	46.327
43.68	66.685	61.483	55.627	49.085
45.68	73.947	68.647	61.903	58.612
48.03	83.325	77.377	69.585	61.328
49.28	89.229	82.094	74.308	65.635
50.75	95.492	88.772	80.292	71.065
52.75	105.287	97.675	88.390	78.287
54.28	113.333	104.379	94.934	83.679
56.05	123.544	114.690	103.942	92.219
57.65	133.313	123.764	112.059	99.404
58.66	139.630	129.432	117.335	104.294
60.28	150.553	139.357	126.410	112.267
62.40	166.218	154.214	139.940	120.683
64.30	181.181	167.809	151.825	135.863
65.42	190.345	176.683	160.395	142.776
67.60	210.000	195.032	176.847	157.844
69.36	226.165	209.611	190.560	169.764
70.65	239.788	222.575	202.610	-
72.00	254.073	235.484	214.240	-
73.40	269.727	250.070	227.763	-
75.43	293.814	272.392	248.017	-
77.15	315.581	292.281	265.725	-
79.33	346.206	321.096	292.914	-
81.16	371.709	344.335	313.954	-
82.58	393.517	365.321	332.993	-
83.38	406.253	376.581	343.240	-
84.70	428.016	395.795	361.092	-
85.63	443.827	412.227	375.827	-
87.15	470.655	436.886	398.170	-
89.43	514.215	476.768	335.715	-
91.25	550.935	510.598	466.386	-
93.28	594.596	552.258	504.084	-
95.43	643.890	597.979	546.451	-
97.20	687.020	634.037	579.670	-
99.50	747,500	693.350	633.205	-

Tammann, 1885

%	p	%	p	%	p
			100°		
1.72	749.8	18.30	601.4	27.96	453.7
2.90	742.5	18.88	579.5	31.05	396.2
7.23	712.8	21.83	553.7	36.66	307.6
10.85	681.3	24.62	511.2	42.81	217.0
12.54	665.5				

Dieterici, 1898

M	p	M	p
	0°		
0	4.620	6.81	3.118
1.355	4.429	8.795	2.516
2.945	4.144	11.86	1.624
5.125	3.598	14.90	1.027

Seelig, 1920

t	p			
	0%	5%	10%	15%
0	4.6	4.5	4.3	3.9
5	6.5	6.4	6.0	5.6
10	9.2	9.0	8.6	7.8
15	12.8	12.4	11.9	10.8
20	17.5	17.2	16.4	14.9
25	23.8	23.3	22.2	20.2
	20%	25%	30%	50%
0	3.3	2.9	2.3	-
5	4.8	4.1	3.3	-
10	6.9	5.8	4.7	0.5
15	9.6	8.0	6.4	1.1
20	13.2	10.9	8.8	2.0
25	17.8	14.9	12.0	2.9

Boswell and Cantelo, 1922

M	Dp/p_0	M	Dp/p_0
	23°		
16.300	0.864	4.200	0.155
12.500	0.700	2.100	0.067
10.100	0.545	0.980	0.025
8.500	0.425		

p_0 = vapour pressure of pure water

Fricke, 1929

mol %	p	
	0°	15°
29.07	3.20	11.38
24.63	5.90	13.55
20.24	10.98	33.92
18.01	15.03	44.74
15.52	20.29	58.45
13.17	25.72	73.20
12.23	28.20	79.59
10.56	31.97	89.35
9.34	34.39	96.05

Osaka, 1929

t	p
sat.sol.	
25	1.72
27	1.91
29	2.13
31	2.39

Hayward and Perman, 1931

%	p	%	p
		30°	
12.84	26.80	38.32	7.97
18.87	23.88	43.02	5.42
26.28	17.85	49.91	3.16
33.28	12.10		
		45°	
10.03	64.65	38.16	19.67
20.81	50.18	44.28	12.35
25.37	42.50	51.21	7.78
33.70	27.14	56.43	5.106
		60°	
8.993	136.4	38.90	40.69
9.963	134.8	42.67	31.70
17.99	114.7	46.05	23.75
19.84	111.5	47.57	21.61
23.93	96.5	53.13	14.15
27.14	86.42	56.19	11.52
29.49	74.00	60.95	9.76
31.86	65.32	64.71	6.32
38.37	43.87		
		70.4°	
0	234.3	39.77	66.82
13.60	193.0	44.52	46.76
21.13	162.4	52.54	26.20
26.93	131.7	62.52	12.99
33.33	101.6		
		80°	
0	355.1	41.66	94.60
10.95	314.1	46.29	69.43
19.90	261.4	51.36	48.54
26.88	209.8	52.35	45.81
32.64	163.5	61.93	22.64
37.81	119.0	66.18	16.41

Shibata, 1935

p	t	p	t
1.72	28	2.13	29
1.91	27	2.39	31

sat. sol. NaOH + 1 aq.

Galinker and Korobkov, 1951

mol %	P kg.	mol %	P kg.
	300°		
5	74.5	55	23.0
20	53.7	60	18.2
30	43.4	70	11.0
40	34.0	75	7.7
50	25.9	80	4.5
	350°		
75	37.0	130	13.3
80	35.0	140	10.0
100	25.5	150	6.0
110	23.5		
	374°		
15	80.0	120	30.0
50	58.5	135	25.3
75	46.6	170	14.3
100	37.0	185	10.0
105	35.0		
	420°		
20	80.5	140	30.8
45	61.0	160	26.0
90	47.0	180	22.0
120	37.0	210	17.5
	460°		
20	80.5	140	30.8
45	61.0	160	26.0
90	47.0	180	22.0
120	37.0	210	17.5

Osaka, 1929

t	p
dissoc. (1+1)	
25	0.199
30	.29
35	.41
40	.50
45	.82

Stokes, 1945

Isopiestic solutions.

m_a	m_b	m_a	m_b	m_a	m_b
		25°			
1.988	1.673	2.115	1.772	2.322	1.928
3.240	2.589	3.865	3.040	4.000	3.141
4.382	3.425	4.647	3.619	5.682	4.376
6.358	4.881	6.509	5.002	6.693	5.144
7.838	6.056	8.803	6.865	10.037	7.937
10.927	8.744	12.551	10.228	13.621	11.240
15.601	13.029	18.140	15.082	19.596	16.184
19.642	16.194	21.926	17.831	23.872	18.990
23.896	19.012	25.826	20.103	26.316	20.346
26.693	20.553	26.802	20.601	26.912	20.691
27.262	20.880	27.447	20.950	28.745	21.650

m = molality : a = sodium hydroxide
b = sulfuric acid

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.925	1.6	0.991
0.2	0.925	1.8	1.002
0.3	0.929	2.0	1.015
0.4	0.933	2.5	1.054
0.5	0.937	3.0	1.094
0.6	0.941	3.5	1.139
0.7	0.945	4.0	1.195
0.8	0.949	4.5	1.255
0.9	0.953	5.0	1.314
1.0	0.958	5.5	1.374
1.2	0:969	6.0	1.434
1.4	0.980		

Gerlach, 1886

%	b.t.	%	b.t.
0	100	80.97	210
14.53	105	82.62	215
23.08	110	84.03	220
29.08	115	85.36	225
33.77	120	87.72	230
37.54	125	88.89	235
41.21	130	89.89	240
44.78	135	90.91	245
48.33	140	91.95	250
51.58	145	93.02	255
54.63	150	93.88	260
57.36	155	94.56	265
60.13	160	95.24	270
62.79	165	95.92	275
65.16	170	96.62	280
67.57	175	97.28	285
69.70	180	97.89	290
71.79	185	98.48	300
73.80	190	99.05	305
75.75	195	99.55	310
77.53	200	sat.sol.	314
79.25	205		

Freezing curve

Rüdorff, 1862

%	f.t.	%	f.t.
1.31	-1.15	6.59	-6.25
2.04	-1.8	7.01	-6.85
4.71	-4.3	9.40	-9.6

Pickering, 1893

%	f.t.	%	f.t.
	NaOH		
83.37	192	75.83	80
81.09	159	74.15	60
78.51	110		
	NaOH + 1 aq.		
72.06	63.5	62.98	58.4
71.59	64.5	62.85	57.8
71.17	63	62.13	55.5
70.76	64.3	62.06	55.5
69.85	65.5	61.31	55.1
69.37	64.4	60.55	53.0
68.45	64	59.31	50.2
67.92	64.2	59.20	49
67.67	63.3	57.95	45.75
67.45	64.5	57.76	44.65
66.85	63.23	56.44	40.25
66.37	63	55.13	34.75
66.04	62.25	51.70	18
64.87	61.5	50.99	13.5
64.61	61.05	50.34	9.0
64.61	61	49.70	4.5
64.27	61		
	NaOH + 2 aq.		
51.70	12.7	46.22	6.7
50.72	11.7	46.06	4.6
49.95	10.7	46.04	5.0
49.81	11.6	46.01	6.5
49.26	10.9	45.78	4.8
49.24	10.0	45.71	5.85
49.11	10.3	45.42	3.9
49.07	11.75	45.39	5.3
49.05	10.35	44.77	2.9
48.59	9.9	44.60	4.35
48.57	11.0	44.60	3.5
48.10	10.4	44.37	2.3
47.95	9.05	44.01	2.1
47.65	7.0	43.95	1.6
47.59	8.0	43.90	1.5
47.30	7.8	43.74	1.3
47.03	6.85	43.54	2.5
47.02	8.7	43.42	1.3
46.76	7.2	42.97	0.35
46.72	6.8	42.77	0
46.52	6.8		
	NaOH + 3 aq.		
43.54	+1.35	40.64	+2.34
43.25	+2.00	40.44	+3.0
42.97	+2.1	40.37	+2.25
42.95	+1.8	40.37	+1.85
42.72	+2.3	40.31	+2.3
42.71	+2.19	40.31	+1.5
42.50	+3.0	40.27	+2.3
42.50	+2.60	39.95	+2.0
42.48	+2.5	39.81	+1.1
42.47	+2.9	39.69	+1.38
42.30	+2.6	39.61	+1.2
42.09	+2.61	39.53	+1.8
41.97	+3.0	39.32	+1.49
41.95	+2.9	39.21	+0.4
41.94	+2.7	39.04	+0.5
41.72	+2.75	38.98	+0.15
41.67	+2.1	38.74	+0.2
41.45	+3.1	38.72	+0.5
41.45	+2.9	38.55	+0.25
41.44	+2.9	38.47	-2.2
41.35	+2.68	38.42	-0.55
41.26	+2.5	38.05	0
41.14	+2.63	37.68	-0.7
41.08	+2.53	37.67	-2.5
41.00	+2.3	37.67	-1.8
40.90	+2.35	37.30	-2.0
	NaOH + 3.5 aq.		
45.52	+7.0	38.22	15.35
44.96	9.3	37.52	15.25
44.80	8.75	37.02	14.5
44.38	11.05	36.65	14.5
44.22	10.75	35.77	13.4
43.76	12.3	35.51	13.2
43.60	12.0	34.70	11.75
43.18	12.3	34.12	10.95
42.96	13.25	34.02	10.1
42.50	14.0	33.68	9.5
42.28	13.95	32.97	7.0
41.86	14.95	32.64	7.03
41.60	14.5	32.39	6.5
40.98	15.2	31.82	3.9
40.80	14.9	31.04	0.3
40.34	15.5	30.84	2.0
39.63	15.5	30.58	-1.1
39.51	15.35	29.58	-6.45
38.99	15.6	29.50	-4.2
38.32	15.6	29.07	-8.75
	NaOH + 4 aq. α		
38.25	+6.6	29.10	- 1.62
37.50	7.08	28.70	- 2.6
37.05	7.3	28.68	- 2.05
36.80	7.4	28.31	- 4.0
36.41	7.47	28.22	- 3.95
36.10	7.55	28.16	- 4.0
35.97	7.6	27.90	- 5.3
35.47	7.55	27.72	- 6.3
35.42	7.54	27.69	- 6.27
34.93	7.4	27.46	- 6.35
34.71	7.35	27.28	- 8.15
34.35	7.2	27.19	- 6.0
34.02	6.8	26.94	- 8.75
34.02	6.7	26.91	- 8.45
33.99	7.0	26.90	- 7.1
33.66	6.83	26.75	-10.5
33.18	6.4	26.50	-12.0
32.92	6.05	26.42	-10.5
32.58	5.95	26.25	-10.3
31.98	5.2	26.20	- 9.85
31.38	4.33	25.94	-12.5
31.29	2.7	25.92	-12.8
31.23	2.4	25.72	-14.1
30.89	3.3	25.58	-13.65
30.82	2.1	25.47	-12.6
30.38	1.65	25.47	-15.58
30.31	1.9	25.46	-14.95
30.21	1.35	25.06	-17.4
29.95	0.7	24.97	-16.75
29.76	1.27	24.95	-17.2
29.53	-0.15	24.64	-18.25
29.14	-1.05	24.33	-19.1
29.13	-2.0	23.80	-21.35

NaOH + 4 aq. β

37.14	-2.25	34.19	- 2.5
36.76	-1.7	34.17	- 3.3
36.64	-1.93	33.78	- 2.85
36.54	-2.15	33.35	- 3.6
36.14	-1.8	33.17	- 2.18
35.93	-1.0	33.06	- 4.35
35.74	-1.8	33.01	- 7.08
35.61	-1.75	32.54	- 5.85
35.45	-1.65	32.49	- 5.7
35.13	-1.9	32.09	- 5.9
35.10	-0.95	31.74	- 4.8
35.03	-1.95	31.51	- 7.08
34.78	-2.1	31.29	- 7.5
34.67	-2.1	31.23	- 7.9
34.24	-2.5	30.18	- 7.95
34.20	-1.5	29.85	-14.45

NaOH + 5 aq.

31.44	-12.75	24.02	-19.9
31.19	-12.6	23.97	-19.55
31.09	-12.5	23.86	-20.1
30.97	-12.47	23.64	-20.8
30.62	-12.28	23.56	-21.0
30.49	-12.25	23.53	-20.4
30.46	-12.25	23.48	-21.45
30.08	-12.20	23.48	-20.8
29.97	-12.29	23.31	-21.7
29.85	-12.15	23.34	-21.7
29.55	-12.42	23.28	-21.0
29.48	-12.25	23.24	-22.2
29.20	-12.35	23.29	-20.8
28.99	-12.48	22.84	-23.6
28.97	-12.50	22.80	-21.75
28.59	-12.50	22.59	-23.4
28.45	-12.78	22.46	-24.0
28.44	-12.80	22.44	-23.7
28.26	-13	22.07	-25.4
27.94	-13.15	22.06	-25.0
27.81	-13.10	22.03	-25.5
27.80	-13.2	21.73	-26.8
27.24	-13.8	21.70	-25.9
26.94	-14.10	21.43	-26.28
26.48	-15.1	21.06	-29.25
26.27	-15.0	21.04	-27.6
25.73	-15.6	21.00	-23.2
25.42	-16.3	20.98	-27.7
25.34	-16.7	20.98	-27.9
25.27	-16.85	20.94	-28.85
25.16	-18.2	20.76	-30.3
24.86	-17.6	20.67	-28.3
24.61	-18.3	20.52	-29.9
24.52	-18.8	20.42	-30.85
24.44	-18.9	20.27	-30.0
24.39	-18.7	20.10	-31.6
24.35	-18.8	20.07	-31.9
24.34	-18.88	19.91	-31.4
24.09	-20.55		

NaOH + 7 aq.

26.91	-24.77	21.53	-24.2
26.25	-24.2	21.48	-25.50
25.59	-23.75	21.40	-24.6
25.44	-23.75	21.33	-25.55
25.02	-23.85	21.10	-25.2
24.87	-23.45	21.02	-24.9
24.42	-25.5	20.98	-25.15
24.20	-23.5	20.97	-25.25
23.75	-23.0	20.94	-25.25
23.70	-23.65	20.64	-25.0
23.56	-23.8	20.62	-25.6
23.22	-23.3	20.59	-25.25
23.02	-23.85	20.52	-25.5
22.93	-23.8	20.43	-26.15
22.86	-23.4	20.30	-25.4
22.63	-23.7	19.98	-26.0
22.62	-24.0	19.92	-26.4
22.32	-24.1	19.84	-27.0
22.30	-24.1	19.80	-25.7
22.22	-24.45	19.60	-26.5
22.13	-24.2	19.68	-26.7
22.06	-24.5	19.46	-26.4
21.85	-24.3	19.44	-27.7
21.56	-24.5	19.39	-26.7
21.56	-24.8	19.00	-28.25
21.55	-24.8		

NaOH + x aq.

20.52	-33.0	9.52	-9.62
19.88	-32.6	9.10	-8.5
19.69	-31.4	8.88	-8.6
19.53	-31.0	8.84	-8.82
19.28	-29.25	7.97	-7.72
19.19	-29.5	7.63	-7.0
18.85	-28.7	7.56	-7.05
18.69	-27.50	6.89	-6.42
18.59	-26.0	6.36	-5.6
18.45	-26.25	6.30	-5.7
18.17	-26.3	5.78	-5.27
18.07	-26.7	5.29	-4.4
17.63	-24.05	5.15	-4.50
17.56	-25.0	4.70	-4.07
16.93	-22.0	4.05	-3.7
16.34	-21.9	3.69	-3.1
16.12	-20.4	3.68	-3.07
15.62	-20.25	3.59	-3.0
15.27	-18.5	2.97	-2.28
14.11	-17.2	2.75	-2.32
14.07	-16.0	2.24	-1.7
14.06	-16.9	2.15	-1.77
12.66	-14.0	1.97	-1.52
12.51	-14.5	1.46	-1.12
11.11	-11.75	1.37	-0.8
10.68	-10.9	0.89	-0.62
10.47	-10.9	0.81	-0.6
10.10	-10.3	0.46	-0.3
10.03	-10.20		

Jones, 1904 and Jones and Bassett, 1905

M	f.t.	M	f.t.
0.05	0.188	1.00	3.465
0.10	0.358	2.00	7.400
0.20	0.693	3.00	11.750
0.30	0.992	4.00	17.000
0.40	1.334	5.00	23.000
0.60	2.030	6.00	33.000
0.80	2.727		

von Antropoff and Sommer, 1926

%	f.t.	tr.t.	%	f.t.
100.0	322	303	91.6	265
98.7	319	305	85.5	205
96.7	298	-		

Properties of phases .

Density .

Tünnermann, 1829

%	d	%	d
15°			
7.800	1.0846	31.202	1.3414
15.601	1.1831	39.002	1.4273
23.401	1.2697		

Schiff, 1858

%	d	%	d
15°			
1	1.011	36	1.394
2	.022	37	.404
3	.034	38	.414
4	.045	39	.425
5	.051	40	.436
6	.069	41	.446
7	.080	42	.455
8	.091	43	.467
9	.102	44	.477
10	.114	45	.488
11	.125	46	.498
12	.136	47	.507
13	.147	48	.518
14	.158	49	.528
15	.169	50	.539
16	.180	51	.549
17	.191	52	.559
18	.201	53	.569
19	.212	54	.579
20	.224	55	.590
21	.235	56	.600
22	.246	57	.610
23	.257	58	.621
24	.268	59	.632
25	.278	60	.642
26	.289	61	.653
27	.299	62	.663
28	.309	63	.673
29	.320	64	.683
30	.331	65	.694
31	.342	66	.704
32	.350	67	.713
33	.362	68	.724
34	.373	69	.735
35	.383	70	.746

Thausen, 1871

mol%	d	mol%	d
18°			
0	0.9987	3.2	1.0782
0.5	1.0124	6.3	.1450
1.0	.0246	11.8	.2576
2.0	.0486		

Berthelot, 1873

mol%	d	mol%	d
14.9°			
0.80	1.023	10.22	1.220
1.40	.035	15.19	.312
1.76	.046	19.64	.383
2.60	.067	23.31	.436
3.49	.088	26.00	.470
6.09	.140	28.01	.494

vanden Willigen, 1874

%	d
21.60°	
11.45	1.12652
18.50	1.20376
34.74	1.37629

Hager, 1876

%	d	%	d
17.5°			
0	0.999	25.80	1.279
7.74	1.082	27.09	.293
9.03	.096	28.38	.307
10.32	.110	29.67	.322
11.61	.124	30.96	.346
12.90	.138	32.25	.351
14.79	.151	33.54	.365
15.48	.165	34.83	.380
16.77	.179	36.12	.394
18.06	.193	38.70	.424
19.35	.208	39.99	.438
20.64	.222	41.28	.453
21.93	.236	42.57	.468
23.22	.250	43.86	.483
24.51	.264	45.15	.498

Kohlrausch, 1879

%	d	%	d
15°			
2.5	1.0280	25	1.2823
5	.0568	30	.3374
10	.1131	35	.3907
15	.1700	40	.4421
20	.2262	42	.4615

Grotrian, 1879

t	d	t	d
19.37%		37.72%	
15.00	1.2147	15.00	1.4161
17.76	.2132	15.44	.4158
18.00	.2131	16.45	.4151
29.32	.2070	29.44	.4063
29.47	.2069	30.31	.4057
38.37	.2018	38.59	.4000
38.71	.2016	39.54	.3993

Pickering, 1894

%	d	%	d
15°			
0	0.999180	25	1.277063
1	1.010611	26	.287990
2	.021920	27	.298877
3	.033109	28	.309708
4	.044317	29	.320496
5	.055463	30	.331213
6	.066602	31	.341879
7	.077733	32	.352472
8	.088856	33	.362991
9	.099969	34	.373453
10	.111069	35	.383815
11	.122165	36	.394092
12	.133250	37	.404279
13	.144353	38	.414363
14	.155450	39	.424353
15	.166538	40	.434299
16	.177619	41	.444161
17	.188707	42	.453929
18	.199783	43	.463623
19	.210861	44	.473249
20	.221933	45	.482850
21	.233062	46	.492406
22	.244119	47	.501927
23	.255134	48	.511412
24	.266092	49	.520868
25	.277063	50	.530282

Moore, 1895-96

M	d	M	d
18°			
0.00	0.9987	2.0	1.0843
0.25	1.0099	4.0	.1551
0.5	1.0212	8.0	.2786
1.0	1.0425		

Bousfield and Lowry, 1905

%	d 0°	10°	20°	30°
0	0.9999	0.9997	0.9983	0.9958
1	1.0124	1.0115	1.0095	1.0069
2	.0244	.0230	.0207	.0177
3	.0364	.0345	.0318	.0285
4	.0482	.0459	.0428	.0393
5	.0598	.0571	.0538	.0501
6	.0713	.0683	.0648	.0609
7	.0828	.0795	.0758	.0717
8	.0943	.0908	.0869	.0826
9	.1057	.1020	.0979	.0934
10	.1171	.1132	.1089	.1043
12	.1399	.1355	.1309	.1261
14	.1624	.1578	.1530	.1480
16	.1849	.1801	.1751	.1699
18	.2073	.2023	.1972	.1918
20	.2296	.2244	.2191	.2136
22	.2519	.2465	.2411	.2454
24	.2741	.2686	.2629	.2571
26	.2963	.2906	.2848	.2789
28	.3182	.3124	.3064	.3002
30	.3400	.3340	.3279	.3217
32	.3614	.3552	.3490	.3427
34	.3823	.3760	.3696	.3632
36	.4030	.3965	.3900	.3835
38	.4234	.4168	.4101	.4035
40	.4435	.4367	.4300	.4232
42	.4632	.4561	.4494	.4425
44	.4825	.4755	.4685	.4615
46	.5018	.4947	.4873	.4805
48	.5210	.5138	.5065	.4994
50	.5400	.5326	.5253	.5181
%	40°	50°	60°	70°
0	0.9924	0.9882	0.9833	0.9586
1	1.0033	0.9990	0.9941	0.9884
2	.0139	1.0095	1.0045	0.9989
3	.0246	.0201	.0150	1.0094
4	.0352	.0305	.0254	.0198
5	.0458	.0412	.0359	.0302
6	.0564	.0517	.0463	.0407
7	.0672	.0623	.0569	.0513
8	.0780	.0730	.0676	.0619
9	.0887	.0836	.0782	.0725
10	.0995	.0942	.0889	.0831
12	.1210	.1157	.1101	.1043
14	.1428	.1373	.1316	.1257
16	.1645	.1588	.1531	.1471
18	.1963	.1805	.1746	.1685
20	.2079	.2020	.1960	.1898
22	.2296	.2236	.2174	.2111
24	.2512	.2451	.2388	.2324
26	.2728	.2666	.2603	.2538
28	.2942	.2878	.2814	.2750
30	.3154	.3090	.3025	.2959
32	.3362	.3298	.3232	.3165
34	.3566	.3501	.3434	.3367
36	.3768	.3702	.3634	.3567
38	.3967	.3900	.3832	.3763
40	.4164	.4095	.4027	.3958
42	.4356	.4287	.4217	.4148
44	.4545	.4475	.4405	.4335
46	.4734	.4663	.4593	.4523
48	.4922	.4851	.4781	.4711
50	.5109	.5038	.4967	.4897

%	80°	90°	100°
0	0.9719	0.9655	0.9586
1	.9824	.9760	.9693
2	.9929	.9865	.9797
3	1.0035	.9970	.9903
4	.0139	1.0075	1.0009
5	.0243	.0179	.0115
6	.0347	.0284	.0220
7	.0453	.0390	.0326
8	.0560	.0497	.0432
9	.0665	.0602	.0537
10	.0771	.0708	.0643
12	.0983	.0920	.0855
14	.1195	.1132	.1066
16	.1408	.1343	.1277
18	.1621	.1556	.1489
20	.1833	.1768	.1700
22	.2046	.1980	.1912
24	.2259	.2192	.2124
26	.2472	.2405	.2336
28	.2682	.2615	.2546
30	.2892	.2824	.2755
32	.3097	.3029	.2960
34	.3299	.3230	.3161
36	.3498	.3429	.3360
38	.3695	.3626	.3556
40	.3889	.3820	.3750
42	.4079	.4009	.3940
44	.4266	.4196	.4127
46	.4454	.4384	.4315
48	.4641	.4572	.4503
50	.4827	.4759	.4690

%	d		%	d	
	18°	15°		18°	15°
0	0.99866	0.99918	26	1.28600	1.2877
1	1.01003	1.01065	27	.2968	.2986
2	.02127	.02198	28	.3076	.3094
3	.03241	.03322	29	.3184	.3202
4	.04349	.04441	30	.3290	.3309
5	.05454	.05554	31	.3396	.3405
6	.06559	.06666	32	.3502	.3624
7	.07664	.07777	33	.3605	.3624
8	.08769	.08887	34	.3708	.3728
9	.09872	.09997	35	.3811	.3830
10	.10977	.11117	36	.3913	.3933
11	.12082	.12217	37	.4014	.4034
12	.13188	.13327	38	.4115	.4135
13	.14294	.14436	39	.4215	.4235
14	.15400	.15545	40	.4314	.4334
15	.16505	.16653	41	.4411	.4432
16	.17610	.17761	42	.4508	.4529
17	.18714	.18868	43	.4604	.4025
18	.19817	.19973	44	.4699	.4720
19	.20920	.21079	45	.4794	.4815
20	.22022	.22183	46	.4890	.4911
21	.23121	.23285	47	.4985	.5007
22	.24220	.24386	48	.5080	.5102
23	.25317	.25485	49	.5174	.5196
24	.26412	.26582	50	.5268	.5290
25	.27506	.27679			

Jones, 1904 and Jones and Bassett, 1905

M	d	M	d
0.05	1.002200	2.00	1.079500
0.10	.004000	3.00	.115700
0.20	.008500	4.00	.153200
0.30	.012500	5.00	.182600
0.40	.016500	6.00	.212900
0.60	.024700	7.00	.243600
0.80	.032800	8.00	.272100
1.00	.040700		

Wegscheider and Walter, 1905 and 1906

%	d	%	d
60°		80°	
22.57	1.2312	22.81	1.2207
20.04	1.2026	14.01	1.1232
17.04	1.1692	0	0.9718
14.16	1.1374		
10.92	1.1020		
0	0.9002		

Chèneveau, 1907

%	d	%	d
18°			
0	0.9986	7.73	1.0864
3.24	1.0360	9.14	1.1018
6.28	1.0701	11.84	1.1320
		14.39	1.1607

Tucker, 1915

mol%	d	mol%	d
17°			
28.87	1.5028	15.00	1.3100
25.85	.4624	8.149	1.1460
23.27	.4344	0	0.9988
19.72	.3844		

Carstens, 1924

%	d	%	d
20°			
0	0.9982	12.04	1.1323
2.74	1.0298	16.66	.1834
5.66	.0618	22.11	.2434
7.40	.0810	27.85	.3064
10.11	.1110	35.00	.3811

Main, 1926

M	d
25°	
1.114	1.0464
1.657	1.0687
2.469	1.0959
3.782	1.1423

Faust, 1927

%	d
20°	
1.02	1.011
5.00	1.0538
10.00	1.1089
20.00	1.2191

Steyer, 1931

M	t	d
0.4982	19.8	1.02040
1.0165	"	.04180
2.1356	"	.08605
3.0516	20.4	.12015
3.6115	19.4	.14020
4.1895	19.2	.15870
5.0460	21.8	.18730
5.2050	19.6	.19285

Hayward and Perman, 1931

%	d	%	d
30°			
12.84	1.1329	38.32	1.4016
18.87	.1988	43.02	.4492
26.28	.2782	49.91	.5124
33.28	.3537		
45°			
10.03	1.0985	38.16	1.3939
20.81	.2128	44.28	.4415
25.37	.2614	51.21	.5104
33.70	.3494	56.43	.5534
60°			
8.993	1.0804	38.90	1.3960
9.963	.0900	42.67	.4320
17.99	.1780	46.05	.4626
19.84	.1946	47.57	.4795
23.93	.2410	53.13	.5260
27.14	.2728	56.19	.5480
29.49	.2990	60.95	.5680
31.86	.3284	64.71	.6010
38.37	.3891		
70.4°			
0	0.97781	39.77	1.3937
13.60	1.1269	44.52	.4380
21.13	.2076	52.54	.5113
26.93	.2683	62.52	.5933
33.33	.3289		
80°			
0	0.97183	41.66	1.4070
10.95	1.0864	46.29	.4494
19.90	.1815	51.36	.4958
26.88	.2562	52.35	.5045
32.64	.3162	61.93	.5869
37.81	.3695	66.18	.6214

Akerlöf and Kegeles, 1939

%	d			
	0°	10°	20°	30°
0	0.99962	0.99948	0.99800	0.99553
2	1.02376	1.02242	1.02022	1.01726
4	.04708	.04489	.04207	.03868
6	.07019	.06726	.06390	.06010
8	.09321	.08963	.08577	.08161
10	.11614	.11120	.10773	.10324
12	.13900	.13445	.12979	.12499
14	.16170	.15686	.15192	.14686
16	.18433	.17925	.17409	.16882
18	.20661	.20155	.19628	.19984
20	.22863	.22371	.21844	.21288
22	.25023	.24567	.24049	.23488
24	.27133	.26731	.26237	.25679
26	.29180	.28855	.28402	.27852
28	.31150	.30929	.30530	.30000
30	.33383	.32940	.32612	.32112
32	.35683	.34996	.34637	.34177
34	.37958	.37243	.36540	.36182
36	.40198	.39654	.38748	.38050
38	.42388	.41629	.40906	.40198
40	.44515	.43744	.43215	.42298
42	.46561	.45738	.45047	.44334
44	.48505	.47732	.46998	.46292
46	.50330	.49468	.48847	.48153
48	.52008	.51272	.50569	.49893
50	.53811	.52814	.52142	.51495
52	.54827	.54174	.53541	.52931

%	40°	50°	60°	70°
0	0.99219	0.98809	0.98330	0.97789
2	1.01360	1.00929	1.00434	.99876
4	.03473	.03021	.02511	1.01939
6	.05587	.05116	.04952	.04006
8	.07711	.07221	.06682	.06085
10	.09849	.09340	.08787	.08179
12	.12000	.11472	.10906	.10289
14	.14164	.13618	.13040	.12416
16	.16340	.15774	.15185	.14555
18	.18523	.17942	.17340	.16706
20	.20711	.20114	.19501	.18866
22	.22898	.22286	.21662	.21029
24	.25078	.24452	.23520	.23189
26	.27244	.26606	.25965	.25342
28	.29389	.28740	.28093	.27479
30	.31503	.30845	.30192	.29592
32	.33576	.32920	.32253	.31672
34	.35595	.34924	.34267	.33705
36	.37547	.36873	.36216	.35682
38	.39485	.38753	.38089	.37587
40	.41585	.40853	.40078	.39403
42	.43625	.43117	.42138	.41318
44	.45591	.44881	.44132	.43336
46	.47465	.46772	.46070	.45284
48	.49226	.48559	.47864	.47137
50	.50855	.50219	.49561	.48882
52	.52326	.51727	.51116	.50492

Hitchcock and Mc Ilkenny, 1935

N	d	N	d
20°			
1.109	1.0454	5.800	1.2090
1.888	.0758	6.032	.2163
2.990	.1165	8.309	.2824
4.677	.1738		
30°			
0.9983	1.0372	5.977	1.2102
1.085	.0407	6.023	.2116
3.018	.1128	8.274	.2771
4.585	.1661		
40°			
0.9751	1.0318	6.161	1.2105
2.010	.0711	6.376	.2172
3.063	.1084	7.883	.2626
3.949	.1387		

Krings, 1947-48

t	d	t	d	t	d
12.6%		23.5%		31.5%	
19.0	1.139	17.6	1.2595	19.4	1.3485
32.5	.133	33.0	.251	30.4	.340
40.8	.133	39.3	.2465	38.7	.3355
49.8	.128	50.2	.241	50.6	.328
60.7	.1145	61.0	.2335	60.3	.321
69.2	.111	69.0	.228	69.2	.316
42.2%		47.7%		56.0%	
20.6	1.3485	23.8	1.508	48.4	1.568
30.5	.447	30.1	.503	50.4	.567
39.5	.441	39.6	.495	60.0	.558
50.7	.433	51.3	.488	70.0	.551
60.5	.426	70.0	.474		
68.5	.420				
58.5%		63.6%		72.8%	
50.0	1.590	65.0	1.6355	70.0	1.689
		70.0	1.629		

Carstens, 1924

%	π	%	π
20°			
0	49.1	12.04	30.0
2.74	43.4	16.66	25.0
5.66	38.4	22.11	20.7
7.40	35.8	27.85	16.8
10.11	32.3	35.00	14.4

Viscosity and surface tension

Grotrian, 1879

t	η	t	η
19.37%		37.72%	
17.76	4652	15.44	42650
18.00	4621	16.45	41530
29.32	3228	29.44	18820
29.47	3215	30.31	18060
38.37	2532	38.59	12480
38.71	2506	39.54	11810

%	η
18°	
0	1103
19.37	4617
37.72	36595

Moore, 1895-96

M	η	M	η
18°			
0.00	1058	2.0	1627
0.25	1121	4.0	3002
0.5	1172	8.0	7995
1.0	1305		

Bousfield and Lowry, 1905

%	η	%	η
18°			
0.0	1000	30.2	11810
2.5	1080	35.0	20600
7.68	1430	40.0	32300
14.28	2250	45.0	48200
20.14	3840	50.6	74700
25.0	6690		

Main, 1926

M	η	M	η
25°			
1.114	1250	2.469	1658
1.657	1395	3.782	2219

Faust, 1927

%	η	%	η
	20°		
1.02	1053	10.00	1870
5.00	1337	20.00	4502

Giordani and Maresca, 1929

t	η	t	η
	N		
11.38	1549	45.25	727
30.00	978	61.18	563
	2 N		
10.95	1932	45.25	874
30.00	1192	61.08	679
	4.08 N		
10.98	3357	45.35	1385
30.00	1962	61.18	1028
202 g/l		250 g/l	
21.7	4205	25.2	3078
35.4	2768	35.4	2103
45.0	2170	44.8	1701

N at 20°

Hitschcock and Mc Ilkenny, 1935

M	η	M	η
	20°		
1.109	1262.1	5.800	4335.1
1.888	1506.7	6.032	4413.2
2.990	1960.4	8.309	8551.1
4.677	3050.8		
	30°		
0.9983	987	5.977	3205
1.085	1001	6.023	3329
3.018	1546	8.274	5773
4.585	2285		
	40°		
0.9751	804.3	6.161	2564.9
2.010	995.6	6.376	2695.8
3.063	1249.8	7.883	3870.0
3.949	1520.9		

Krings, 1947 - 1948

%	η			
	22°	30°	35°	40°
0	954	817	731	672
12.6	2060	1690	1530	1040
23.5	6000	4500	3750	3250
31.5	14670	10070	8320	6910
42.2	41600	25130	18680	14800
47.7	58900	35000	26900	20800
	48°	50°	55°	
0	578	557	530	
12.6	1210	1160	1080	
23.5	2060	2450	2130	
31.5	5230	4880	4200	
42.2	10620	9770	8180	
47.7	14230	12950	10500	
56.0	21800	19700	16000	
	60°	65°	70°	
0	469	435	411	
12.6	990	910	840	
23.5	1950	1760	1570	
31.5	3620	3200	2850	
42.2	6960	6050	5350	
47.7	8850	7550	6500	
56.0	13200	11000	9170	
63.7	-	16800	13800	
72.8	-	-	21500	

Faust, 1927

%	σ
20°	
1.02	73.4
5.00	74.8
10.00	77.0
20.00	85.0

Optical and electrical properties .

van der Willigen, 1874

spectral lines	n 11.45%	18.50%	34.74%
		21.60°	
A	1.35898	1.37543	1.40757
a	.35995	.37647	.40875
B	.36073	.37731	.40968
C	.36153	.37816	.41071
D	.36365	.38044	.41334
E	.36625	.38323	.41651
b	.36675	.38375	.41719
F	.36847	.38560	.41936
G	.37069	.38800	.42218
G	.37248	.38990	.42441
H	.37319	.39071	.42537
H	.37438	.39203	.42691
H	.37578	.39358	.42872

Briner, 1906

g/l	n_D	g/l	n_D
	18°		
36.87	1.3479	60.00	1.3679
42.86	.3519	60.71	.3682
50.00	.3574	63.64	.3717
50.69	.3579	66.05	.3753
55.55	.3625		

Cheneveau, 1907

%	n_D	%	n_D
	18°		
0	1.3331	9.14	1.3574
3.24	.3419	11.84	.3646
6.28	.3501	14.39	.3711
7.73	.3539		

Steyer, 1931

M	n C	D	F
		20°	
0.4982	1.33661	1.33848	1.33273
1.0165	.34173	.34367	.34802
2.1356	.35223	.35425	.35883
3.0516	.36441	.36656	.37142
4.1895	.36836	.37056	.37550
5.0460	.37491	.37712	.38218
5.2050	.37587	.37809	.38323

Kohlrausch, 1879

%	ϰ	$\tau.10^4$	%	ϰ	$\tau.10^4$
		18°			
2.5	1083	195	25	2703	370
5	1962	202	30	2011	452
10	3111	218	35	1498	554
15	3448	250	40	1158	652
20	3255	301	42	1058	695

Kunz, 1903

t	ϰ	t	ϰ
27.11 %		32.70 %	
0	1030.65	0	682.40
-10.09	630.94	-10.96	327.81
-20.16	341.78	-20.25	109.69
-30.19	107.41		

Bachofner, 1904

t	ϰ 0.5 M	1.0 M	1.5 M	2.0 M	2.5 M	3.0 M
			18°			
18	86	1600	2110	2540	2880	3130
20	89	1660	2200	2650	3010	3280
25	98	1810	2410	2920	3340	3650
30	107	1950	2630	3200	3660	4030
35	115	2100	2850	3480	3990	4400
40	124	2260	3070	3760	4330	4770

Jones, 1904 and Jones and Bassett, 1905

M	λ	M	λ
0.05	121	2.00	78.50
0.10	119	3.00	62.30
0.20	117	4.00	49.67
0.30	113	5.00	38.27
0.40	110	6.00	29.28
0.60	105	7.00	21.67
0.80	101	8.00	16.07
1.00	95.50		

Bousfield and Lowry, 1905

%	ϰ 0°	18°	50°	100°
49.68	177.5	830	3918	12830
46.74	216.8	923.5	4081	12940
43.56	265.6	1038.4	4286	13190
40.40	334.9	1181	4503	13400
39.72	359.0	1224.5	4559	13480
34.43	572.8	1613	5157	13850
29.76	887.3	2114	5798	14080
26.68	1147	2494	6255	14110
19.96	1810	3272	6776	13250
14.96	2106	3493	6505	11590
10.56	2020	3175	5601	9246
3.85	1047	1575	2600	4068

Demolis, 1906

t	κ 58.44 g/l	102.8 g/l	154.5 g/l	194.6g/l
20	2158	3060	3556	3584
30	2570	3727	4390	4588
40	3015	4410	5272	5569
50	3456	5099	6181	6586
60	3889	5794	7100	7655
70	4313	6470	8026	8783
80	4737	7151	8974	9897

g/l	κ 18°	25°	50°	75°
50	1871	2142	-	-
75	2428	2785	4213	5455
100	2899	3342	5027	6698
125	3213	3699	5641	7669
150	2875	3942	6112	8383
175	3420	4070	6412	8968
200	3342	4099	6619	9425

Giordani and Maresca, 1929

t	κ	t	κ
	N (20°)		
11.28	1387	44.65	2451
30.00	1967	61.08	2985
	2 N		
11.28	2306	45.30	4165
20.00	3307	61.08	5446
	4.08 N		
11.23	2973	45.65	5973
30.00	4544	60.78	7492
202 g/l		250 g/l	
12.5	2965	11.6	2660
34.6	5027	34.7	4914
45.3	6152	45.3	6130
63,2	8076	62.8	8278

Sloan, 1910

M	λ	M	λ
	0°		
0.531	36.41	4.250	51.50
1.0625	43.20	8.50	54.72
2.125	47.81	17.00	57.41

Dewar and Fleming, 1897

t	ε	t	ε
sat.sol.		10%	
-201.8	7.02	-199.8	50.6
-181.8	7.08	-175.0	76.7
-146.7	7.70	-156.0	117.2
-102.2	10.20	-147.2	137.0
-91.8	23.10		

Okazaki, 1933

%	Verdet's constant (3441 Å)	%	Verdet's constant (3441 Å)
	28°		
4.03	0.04637	36.89	0.06706
9.08	.04954	36.94	.06672
17.53	.05458	38.55	.06771
22.36	.05775	42.64	.07053
26.03	.06003	44.20	.07195
29.77	.06248	44.29	.07206
33.16	.06489		

Heat constant

Thomsen, 1871

mol %	U	mol %	U
	18°		
0	1.000	3.2	0.919
0.5	0.983	6.3	0.878
1.0	0.968	11.8	0.847
2.0	0.942		

Hammerl, 1879

%	U	%	U
	(12°-30°)		
0	1	25.55	0.869
9.29	0.924	32.92	0.852
18.56	0.886	49.42	0.816

Blümcke, 1885

%	U
53	0.81
61	0.85
73	0.96

Tucker, 1915

mol %	U	mol %	U
	20°		
0	1.000	23.27	0.785
8.149	0.844	25.85	0.784
15.00	0.807	28.87	0.783
19.72	0.790		

Richards and Rowe, 1921

mol % initial	final	U
	20°	
9	3.8	0.9046
9	2	0.9418
9	1	0.9670
9	0.5	0.9818
9	0.25	0.9907
14.8	1	0.9668
14.8	0.5	0.9828

Bertetti and Mc Cabe, 1936

%	t	U
4.081	3.02	0.9380
	8.27	.9416
	20.00	.9476
	30.00	.9523
5.78	42.10	0.9406
	50.28	.9428
	60.49	.9454
	67.93	.9476
6.80	19.89	0.9230
6.920	4.15	0.9113
	7.16	.9141
	13.27	.9184
	24.96	.9254
7.39	19.82	0.9179
	20.82	.9184
8.91	19.94	0.9070
11.90	39.44	0.9027
	51.22	.9046
	59.57	.9086
	68.15	.9108
17.70	20.26	0.8695
	26.04	.8731
19.89	42.22	0.8776
	49.15	.8802
19.89	60.27	0.8840
	70.21	.8857
20.98	19.70	0.8607
	25.60	.8645
	30.00	.8683
21.38	5.96	0.8465
	6.77	.8473
	10.20	.8506
	10.87	.8520
	16.54	.8566
	20.78	.8590
	25.17	.8623
	30.20	.8666
29.53	39.87	0.8580
	50.23	.8633
	60.60	.8676
	68.55	.8688
29.71	8.08	0.8286
	14.56	.8344
31.52	20.18	0.8372
	25.22	.8413
	30.62	.8431
37.10	19.96	0.8248
	25.21	.8286
37.53	5.68	0.8150
	9.68	.8172
	14.68	.8208
40.44	20.16	0.8153
	25.11	.8179
	30.22	.8198
39.44	39.71	0.8286
	40.97	.8294
	49.97	.8320
	60.77	.8337
	69.49	.8348
41.48	20.60	0.8126
	25.11	.8143
	30.53	.8148
45.20	61.18	0.7997
	69.19	.7981
	80.15	.7971
47.71	20.60	0.7857
	24.84	.7840
	30.60	.7849
48.70	65.75	0.7744
	80.18	.7728
	88.53	.7723
50.88	13.32	0.7685
	20.13	.7670
	25.50	.7661
	34.00	.7651
51.15	50.12	0.7606
	60.22	.7596
	68.85	.7587
54.42	27.54	0.8036
	39.26	.8052
	50.90	.8057

Wilsen and Mc Cabe, 1942

t	U	t	U
50.2%		51.5%	
29.748	0.7717	31.35	0.7651
35.54	.7710	51.7%	
50.457	.7664	27.539	0.7521
52.006	.7670	33.29	.7510
72.099	.7640	36.14	.7621
92.769	.7615	55.238	.7550
112.337	.7606	75.39	.7472
104.13	.7623	76.1246	.7534
122.76	.7619	60.3%	
55.3%		71.027	0.6985
48.458	0.7329	73.37	.6980
52.477	.7340	85.836	.6937
54.317	.7341	100.09	.6935
41.437	.7343	114.335	.6889
43.083	.7346	116.069	.6838
66.539	.7323		
82.648	.7291		
84.38	.7269		
99.906	.7252		
117.95	.7217		
119.565	.7253		
64.1%		71.1%	
64.025	0.6812	74.11	0.6475
71.71	.6793	75.865	.6463
73.77	.6807	80.35	.6500
85.65	.6790	82.30	.6459
95.16	.6759	90.15	.6449
96.82	.6763	91.94	.6436
106.94	.6735	100.89	.6423
108.529	.6815	102.70	.6428
110.11	.6748	111.28	.6413
115.027	.6689		
65.6%		75.9%	
72.133	0.6738	113.21	0.6242
73.845	.6741	123.825	.6225
80.63	.6716		
82.40	.6704		
89.74	.6701		
100.079	.6680		
101.989	.6661		
115.957	.6610		
117.867	.6650		

Tucker, 1915

mol% initial	mol% final	Q dil. (by mole aq.)
27.18	25.36	1838
23.46	21.69	1299
21.69	20.09	1069
20.09	18.67	919
18.67	17.24	727.4
17.24	15.93	539
15.93	14.73	414.9
14.73	13.66	303.0
13.66	12.62	234.2
12.62	11.61	169.2
10.15	8.75	37.8

Richards and Rowe, 1921

mol% initial	mol% final	Q dil.
	20°	
9	3.8	+56
"	2	-55
"	1	-127
"	0.5	-150
"	0.25	-161
14.8	1	+564
"	0.5	+541

Berthelot, 1873

mol% initial	mol% final	Q dil. (by mole NaOH)
28.01	0.22	+3.69
26.00	1.16	+3.18
23.31	.53	+2.41
19.64	.31	+1.47
15.19	.66	+0.38
10.22	2.12	-0.20
6.09	1.28	-0.29
5.15	1.61	-0.39
3.49	3.49	-0.24
2.60	1.33	-0.245
1.76	1.75	-0.145
1.40	0.78	-0.155
1.23	1.22	-0.075
0.80	0.80	-0.06
0.44	0.44	-0.02

Berttetti and Mc Cabe, 1936

% initial	% final	Q dil. (cal/g)
	20°	
48.14	47.12	-5.314
47.12	46.09	5.451
46.09	45.10	5.230
45.10	41.77	17.52
41.77	38.80	15.06
38.80	35.88	13.79
35.88	33.11	11.73
33.11	30.44	9.78
30.44	28.00	7.55
28.00	25.71	5.81
25.71	23.44	4.645
23.44	21.30	3.421
21.30	19.35	2.378
19.35	17.56	1.588
26.14	22.46	7.27
22.46	19.15	4.36
19.15	16.14	2.27
16.14	13.59	0.60

Wilsen and Mc Cabe, 1942

% initial	% final	Q dil
	93.3°	
75.51	74.04	-758
74.42	73.01	745
73.04	70.68	731.7
70.76	68.69	717.8
68.88	66.77	697.1
66.81	64.53	685.0
64.29	62.04	667.4
62.19	60.06	645.5
60.03	57.83	633.1
57.85	55.58	585.4
55.68	53.43	594.2
56.87	54.85	603.9
55.17	52.78	590.1
52.69	50.65	567
50.90	48.68	540
48.64	46.43	514

Water (H_2O) + Sodium hypochlorite (NaClO)

Sanfourche and Gardent, 1924

%	f.t.	E	%	f.t.	E
5	-2.6	-	49.2	26.0	24.0
9.8	-6.5	-16	50.2	30.5	23.0
14.6	-12.0	-16.5	52.0	30.5	23.7
16.6	-12.7	-16.7	53.0	40.5	24.0
17.5	-13.6	-16.7	54.6	42.5	22.3
19.2	-16.6	-16.6	55.4	45.5	22.5
19.7	-13.5	-16.5	57.0	51.5	22.5
20.0	-11.7	-16.4	59.0	54.0	23.5
20.9	-7.0	-16.1	61.2	56	23.5
30.6	+15.0	-17.5	64.3	"	57.5
33.3	19.5	-15.7	66.0	"	57.5
34.8	20.0	-15.7	67.5	"	57.5
37.5	21.0	-17.4	68.5	"	57.8
40.0	25.0	-17.0	70.5	"	59.0
43.3	24.0	"	79.0	"	57.0
44.6	25.5	"	84.0	-	70.0 x
44.8	24.0 (4+1)	"	87.0	-	70.0
45.8	23.7	"	94.0	-	67.0
48.8	25.0	+24.5	x = exploses.		

Water(H_2O) + Sodium borate ($NaBO_2$)

Menzel and Schulz, 1943

t	p	t	p
	4 aq.		
14.4	11.1	20.1	15.9
15.8	12.2	21.7	17.3
17.0	13.0	23.9	19.8
18.8	14.8		
	2 aq.		
19.3	6.5	24.8	10.0
22.0	8.3	27.0	11.9

%	m	f.t.	%	m	f.t.
1.011	0.1551	-0.511	10.74	1.828	-4.684
1.731	0.2676	-0.850	11.15	1.906	-4.632
2.695	0.4208	-1.286	12.02	2.097	-5.233
4.853	0.7749	-2.148	12.61	2.192	-5.524
5.732	0.9238	-2.532	13.01	2.272	-5.724
7.317	1.199	-3.174	13.10	2.289	-5.768
9.511	1.597	-4.084			

E : -5.768

Blasdale and Slansky, 1939

%	f.t.	%	f.t.
14.10	1	37.85	54
15.70	0	38.12	55
17.23	5	39.00	60
18.50	10	40.35	65
20.22	15	41.90	70
22.00	20	43.80	75
23.90	25	45.80	80
26.22	30	48.10	85
28.75	35	50.35	90
31.31	40	52.80	95
35.02	50	55.60	100

tr.t.(4 aq. - 2 aq.) = 54°

Water (H_2O) + Sodium nitrite ($NaNO_2$)

Tammann, 1885

%	p	%	p
		100°	
2.93	749.9	31.86	595.7
5.95	737.9	36.25	566.0
9.63	722.9	41.96	525.8
18.90	676.3	51.36	453.9
24.02	647.3	52.78	441.7

Helberg, 1921

t	%	p 0%	p sol.
21	47.08	18.65	14.02
22	47.2	19.83	14.87
23	47.4	21.07	15.76
24	47.6	22.38	16.70
25	47.9	23.76	17.68

%	f.t.	%	f.t.
9.8	-5.0	20.35	-14
12.65	-7.0	30.8	-20.5
16.1	-10.0	38.0	-26 E
17.35	-11		

Oswald, 1914

%	f.t.	%	f.t.
9.1	-4.5	51.4	52.5
23.8	-9	54.6	65
29.6	-12.5	57.9	81
39.7	-15.5	59.7	92
40.8	-8	62.6	103
41.9	0	68.7	128
44.9	19	100	217

45.8% 20° (sat.sol.) d= 1.3583

b.t.(761.5 mm) = 128°

Kordes, 1926

mol%	f.t.
0	0
14.6	-16 E
100	+282

Bureau, 1934 and 1937

%	f.t.	E	tr.t. (1+2)
5.92	-2.8	-19.5	-
10	-3.8	-19.4	-
16.3	-8.7	-19.5	-
21	-11.5	-19.4	-
25	-15.2	-19.5	-
28.10	-19.5	-	-
28.21	-19.5	-19.5	-
34.5	-12.0	-19.5	-
37.8	-7.6	-19.3	-
41.60	-5.1	-19.5	-
41.65	-5.1	-19.5	-5.1
43.7	+15	-19.5	-4.7
45.07	22	-19.6	-5.0
52.00	56.8	-19.4	-5.1
54.07	64.7	-19.5	-5.1
56.95	79.2	-19.5	-5.1
61.51	99.0	-	-5.1

Traube, 1895

%	d	%	d
	15°		
0	0.9991	8.512	1.05735
1.988	1.01253	19.138	1.13306
4.661	1.03041		

Bureau, 1934 and 1937

%	t	d	%	t	d
43.7	15.0	1.344	54.07	64.7	1.395
45.07	22.0	.347	56.95	79.2	.414
52.00	56.8	.384	61.5	99.0	.448

Water (H_2O) + Sodium chlorate ($NaClO_3$)

Wüllner, 1860

t	p		
	0%	11.11%	20.00%
37.4	48.73	47.58	-
39.75	54.16	52.56	51.26
41.00	57.91	56.16	54.41
43.20	64.20	62.25	60.15
45.70	73.98	71.80	69.62
47.50	81.14	78.95	76.49
49.40	89.21	86.76	84.31
51.20	97.62	94.77	92.22
53.30	108.75	105.50	102.40
55.65	120.89	117.29	113.14
57.45	131.73	127.73	124.23
59.30	144.69	140.59	136.54
62.70	169.25	164.50	159.70
64.25	180.37	175.07	169.77
65.10	187.79	182.39	177.39
67.60	209.89	203.39	197.74
69.40	217.12	210.27	204.27
71.07	244.09	236.69	228.79
73.20	267.42	259.02	251.52
75.30	291.17	283.12	274.47
78.20	329.53	319.13	309.53
81.62	378.34	367.34	356.84
84.50	424.61	412.61	388.61

Tammann, 1885

t	p				
	0%	19.68%	34.70%	43.65%	57.77%
19.11	16.6	15.4	13.8	11.7	-
25.79	24.9	23.0	21.2	17.7	-
31.00	33.7	30.8	27.8	23.7	-
36.22	45.1	41.8	37.8	31.6	-
43.61	66.9	61.6	56.1	46.9	-
47.90	83.3	76.7	69.9	58.4	-
52.39	104.1	95.5	87.1	73.4	-
53.86	111.8	102.6	93.6	78.6	-
57.42	132.5	122.1	110.8	92.8	-
60.54	153.2	140.9	128.2	105.5	99.0
63.11	172.3	159.0	144.4	120.9	113.8
66.09	197.0	181.9	165.4	136.9	126.0
68.53	219.3	202.4	183.5	153.4	141.0
71.26	246.8	227.5	205.9	173.0	158.2
73.88	275.9	255.0	231.4	193.1	176.4
75.28	292.8	269.2	244.9	204.6	186.9
76.86	312.6	288.5	263.6	219.6	200.4
79.11	342.8	316.3	286.9	239.9	219.5
81.28	374.3	345.8	314.4	262.2	240.5
83.52	409.2	378.8	344.1	286.5	261.7
88.46	496.0	459.6	416.1	347.6	317.0
90.79	541.9	500.6	453.9	379.6	347.0
92.87	585.9	539.1	490.3	410.1	374.5
95.48	645.3	591.9	539.7	451.5	412.7
100.13	763.6	709.4	639.4	545.6	-

Tammann, 1885

%	p	%	p
	100°		
8.28	740.3	36.31	627.9
19.30	705.3	46.15	562.6
22.68	692.5	54.80	517.1
31.61	652.7	61.56	470.0

Carlsen, 1932

c	b.t.	c	b.t.
16.5	101.2	99.0	115.0
51.2	106.4	112.0	118.0
81.5	110.7	121.0	121.5

(sat.sol.)

Kremers, 1854

%	f.t.	%	f.t.
45.45	0	64.10	70
49.50	12	71.43	90
55.87	30	75.19	115
60.24	50		

Kremers, 1856

%	f.t.	%	f.t.
45.05	0	59.52	60
49.75	20	63.69	80
55.25	40	67.12	100

b.t. sat.sol. = 132°

Dupuy, 1884

%	f.t.
46.15	15
50.37	30
61.24	60
64.02	100

Billiter, 1902

c	f.t.	c	f.t.
72.2	20	96	70
77	30	100.2	80
82	40	106	90
86.6	50	111	100
91.3	60		

Carlsen, 1910

%	f.t.	%	f.t.
41.86	-15	60.78	60
44.13	0	65.40	80
50.24	20	69.70	100
55.75	40	74.10	122

Le Blanc and Schmandt, 1911

%	f.t.	%	f.t.
45.47	4.78	52.96	35.10
48.91	19.85	54.50	44.72
51.22	30.05		

Cavallaro, 1941 and 1943

m	f.t.	m	f.t.
0.000987	-0.00363	0.50729	-1.6464
.002002	-0.00733	.65031	-2.0985
.005008	-0.01818	.75792	-2.4322
.009996	-0.03598	.87033	-2.7894
.025441	-0.09018	1.03488	-3.3519
.05017	-0.17520	.2523	-4.151
.00996	-0.34326	.5753	-5.403
.20119	-0.67640	.7558	-6.102
.30302	-1.0057	2.1558	-7.590
.42061	-1.3804		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.927	1.6	0.874
0.2	0.913	1.8	0.875
0.3	0.904	2.0	0.876
0.4	0.897	2.5	0.879
0.5	0.892	3.0	0.881
0.6	0.888	3.5	0.886
0.7	0.885	4.0	-
0.8	0.883	4.5	-
0.9	0.882	5.0	-
1.0	0.880	5.5	-
1.2	0.878	6.0	-
1.4	0.876		

Kremers, 1855

%	d	%	d
	19.5°		
1	1.005	21	1.153
2	.013	22	.163
3	.019	23	.172
4	.026	24	.179
5	.033	25	.188
6	.039	26	.198
7	.046	27	.206
8	.053	28	.214
9	.061	29	.223
10	.068	30	.233
11	.076	31	.242
12	.083	32	.250
13	.092	33	.260
14	.098	34	.270
15	.106	35	.280
16	.114	36	.289
17	.122	37	.299
18	.129	38	.309
19	.138	39	.319
20	.145	40	.329

Loewenfeld, 1905

%	d	%	d
	15°		
0	0.9993	18.7	1.1372
5.2	1.0371	26.6	1.2049
10	1.0709		

Carlson, 1910

%	d	%	d
41.86	1.380	60.78	1.514
44.13	.389	65.40	.559
50.24	.400	69.70	.604
55.75	.472	74.10	.654

Heydweiller, 1912

M	d	M	d
	18°		
0.1025	1.00588	1.041	1.07057
.2054	.01315	2.082	.14040
.513	.03467	4.164	.27524

Clausen, 1912

M	d 6°	d 18°	d 30°
0.517	1.03747	1.03483	1.03093
1.0233	.07327	.06944	.06471
2.020	.14174	.13638	.13027
4.012	.27343	.26550	.25754

Lübben, 1914

%	m	d
	18°	
0	-	0.99862
5.79	0.563	1.0385
10.18	1.023	1.0698
18.76	1.1354	1.1354
32.59	1.2557	1.2556

Rubien, 1921

M	d	M	d
	18°		
0	0.99862	1.0360	1.06924
0.1039	1.00607	2.0084	1.13636
0.2061	1.01340	3.987	1.26547
0.5138	1.03484		

Andersen and Asmussen, 1932

%	d
0°	
43.26	1.3771
23.91	.1941

Loewenfeld, 1905

%	σ	%	σ
	15°		
0	75.135	18.7	68.69
5.2	71.99	26.6	68.49
10	70.59		

Lübben, 1914

18°

%	m	n 4600.9 Å	3995.2	3611.9
0	-	1.33878	1.34332	1.34755
5.79	0.563	.34442	.34807	.35341
10.18	1.023	.34886	.35353	.35795
18.76	2.000	.35774	.36275	.36720
32.59	3.845	.37348	.37858	.38344

%	n 3435.3 Å	3345.4	3302.9	3085.1
0	1.35010	1.35159	1.35233	1.35681
5.79	.35601	.35752	.35827	.36284
10.18	.36058	.36214	.36291	.37757
18.76	.36994	.37156	.37243	.38720
32.59	.38633	.38804	.38887	.39405

%	2981.1 Å	2881.1	2801.0	2771.0
0	1.35944	1.35944	1.36224	1.36576
5.79	.36557	.36845	.37102	.37208
10.18	.37034	.37327	.37589	.37694
18.76	.38005	.38310	.38587	.38697
32.59	.39716	.40043	.40336	.40452

%	2743.7 Å	2573.2	2558.1	2502.1
0	1.36653	1.37359	1.37430	1.37708
5.79	.37287	.38011	.38086	.38370
10.18	.37778	.38521	.38597	.38898
18.76	.38782	.39562	.39642	.39952
32.59	.40548	.41380	.41463	.41797

Walter, 1925

15°

%	n_D	%	n_D
0	1.3334	25.5	1.3600
4.4	1.3380	35.8	.3721
10.4	1.3436		

Rubien, 1911

18°

M	n_D	M	n_D
0	1.33327	1.0360	1.34329
0.1039	.33435	2.0084	.35237
0.2061	.33538	3.987	.36905
0.5138	.33845		

Heydweiller, 1913

18°

M	n_D	M	n_D
0	1.33327	1.0	1.34313
0.1	1.33431	2.0	1.35229
0.2	1.33532	4.0	1.36916
0.5	1.33833		

Heydweiller, 1912

18°

M	ϰ	M	ϰ
0.1025	82.4	1.041	636.7
0.2054	156.3	2.082	1060
0.513	353.8	4.164	1461

Clausen, 1912

M	t	λ	$\tau \cdot 10^4$
0.517	6	51.58	266.6
	18	68.73	355.4
	30	87.20	450.8
1.0233	6	46.47	475.5
	18	61.60	630.4
	30	77.92	797.3
2.020	6	39.05	789.2
	18	51.63	1043
	30	65.09	1315
4.012	6	27.12	1088
	18	36.23	1453
	30	46.04	1847

Andersen and Asmussen, 1932

0°

%	Verdet's constant
43.26	0.0174
23.91	0.0164

0°

%	(α) mol magn
43.26	3.35
23.91	3.10

Water (H_2O) + Sodium bromate ($NaBrO_3$)

Tammann, 1885

%	p	%	p
8.63	744.5	35.56	659.6
26.43	695.1	45.72	609.8
29.06	686.1		

van t'Hoff and Armstrong, 1903

t	p	t	p
	sat.sol.		
40	50.7	56	107.9
50.2	83.3	59.5	129.1
55	103.8	60	131.6

Kremers, 1856

%	f.t.	%	f.t.	
21.59	0	43.10	80	
27.71	20	47.62	100	
33.44	40	sat.sol.	109	b.t.
38.46	60			

Robinson and Stokes, 1949

m	Osmotic coefficient	m	Osmotic coefficient
	25°		
0.1	0.918	0.9	0.839
0.2	0.896	1.0	0.833
0.3	0.883	1.2	0.824
0.4	0.873	1.4	0.815
0.5	0.865	1.6	0.808
0.6	0.857	1.8	0.804
0.7	0.851	2.0	0.800
0.8	0.845	2.5	0.792

Kremers, 1855

%	d	%	d	%	d
		19.5°			
1	1.007	11	1.091	21	1.186
2	1.014	12	1.099	22	1.197
3	1.023	13	1.108	23	1.208
4	1.031	14	1.118	24	1.218
5	1.039	15	1.127	25	1.229
6	1.047	16	1.135	26	1.240
7	1.054	17	1.144	27	1.251
8	1.064	18	1.154	28	1.263
9	1.073	19	1.164	29	1.275
10	1.081	20	1.176	30	1.287

Heydweiller, 1921

N	d	λ
	18°	
0.2	1.02338	69.4
0.5	.0581	61.8
1	.1154	54.5
2	.2278	44.06

Ricci, 1934

%	t	d	%	t	d
21.42	5	1.194	29.85	30	1.284
23.24	10	.211	31.35	35	.288
24.94	15	.232	32.80	40	.310
26.69	20	.248	34.22	45	-
28.29	25	.257	35.55	50	-

Water (H_2O) + Sodium iodate ($NaIO_3$)

Kremers, 1856

%	f.t.	%	f.t.	
2.45	0	21.69	80	
8.31	20	25.32	100	
12.58	40	sat.sol.	102	b.t.
17.27	60			

Foote and Vance, 1928, 1929 and 1933

%	f.t.		%	f.t.	
2.42	0.0				
3.90	8.0	5 aq.	15.91	57.8	
4.39	10.0	"	19.03	69.6	
5.88	15.0	"	19.04	67.0	0 aq.
7.84	20.0	1 aq.	19.56	70.6	"
8.66	25.0	"	20.49	75.8	"
9.63	30.0	"	21.24	80.6	"
10.57	35.0	"	21.82	79.0	"
11.71	40.0	"	22.22	87.6	"
14.06	49.9	"	23.03	90.3	"

Cornec and Spack, 1931

%	t		%	t	
2.45	0	5 aq.	16.65	60	1 aq.
4.44	10	"	19.24	70	"
5.93	15	"	22.18	80	"
7.77	20	5 aq. + 1 aq.	21.25	80	$NaIO_3$
9.63	30	1 aq.	22.87	90	"
11.64	40	"	24.70	100	"
13.90	50	"			

Water + Sodium nitrate ($NaNO_3$)

Heterogeneous equilibria.

Wüllner, 1858

t	p 0%	p 0.09%	p 16.67%	p 23.08%
23.1	21.01	20.51	19.82	19.42
27.1	26.66	26.06	25.32	24.77
29.1	29.95	29.16	28.41	27.62
31.5	34.36	33.27	32.28	31.28
37.5	48.25	46.86	45.37	43.88
39.4	53.14	51.46	49.87	48.38
41.5	59.48	57.69	55.76	54.03
44.0	67.79	65.66	63.33	61.22
47.3	80.39	77.56	74.74	71.96
49.1	87.93	85.25	82.18	79.41
51.5	99.10	95.83	92.57	89.20
55.1	118.40	114.19	110.26	106.24
56.5	126.24	122.18	117.93	113.92
58.5	138.76	134.40	130.06	125.50
62.5	166.98	160.85	154.92	149.58
64.5	182.12	175.60	169.08	162.85
65.6	201.07	194.30	187.74	180.42
68.9	222.29	214.69	206.79	199.48
70.5	238.24	229.56	220.97	212.28
72.5	260.60	251.72	242.83	234.24
75.4	293.00	281.76	271.69	261.82
78.7	336.33	324.41	312.57	301.42
81.5	376.85	363.05	349.15	336.15
83.4	406.57	391.50	376.24	361.55
86.0	450.34	433.80	417.35	401.11
88.7	499.98	481.38	467.96	446.83
92.6	579.67	559.21	537.77	517.41
94.9	631.44	609.34	586.74	566.06
100.3	768.20	740.92	714.00	689.13

Pauchon, 1880

t	p	t	p
	41.08%		
7.35	6.10	19.11	12.93
7.98	6.37	20.01	13.57
8.49	6.58	20.92	14.64
8.99	6.78	21.77	15.22
9.52	7.03	22.63	16.03
10.03	7.25	23.57	16.91
10.51	7.62	24.31	17.65
11.17	7.90	25.07	18.48
11.92	8.14	25.98	19.44
12.51	8.66	26.76	20.37
13.05	8.79	27.49	21.28
13.60	9.17	28.34	22.31
14.10	9.47	29.17	23.40
14.76	9.92	30.04	24.39
15.27	10.17	31.00	25.82
15.83	10.55	31.85	27.09
16.38	10.88	32.76	29.54
16.94	11.38	33.45	29.58
17.92	11.85	35.01	31.93
18.28	12.26		

t	p	t	p
	35.49%	23.63	17.35
12.10	8.54	24.18	17.92
12.79	8.93	25.31	19.05
13.51	9.38	25.86	19.78
14.30	9.73	26.16	20.22
15.11	10.27	26.31	20.37
15.92	10.67	26.35	20.37
16.70	11.50	27.63	21.89
17.21	11.86	28.43	22.90
18.11	12.53	29.31	24.16
18.72	12.94	30.01	25.11
19.31	13.38	31.52	27.07
19.83	13.82	31.99	27.79
20.62	14.38	32.41	28.49
21.31	15.11	32.82	29.16
21.93	15.86	33.22	29.79
22.38	15.97	34.68	32.15
22.77	16.52	35.19	32.80
23.02	16.75	38.00	48.71
	23.85%		
8.05	7.18	18.26	13.81
8.31	7.31	18.75	14.28
8.63	7.44	19.11	14.58
8.99	7.62	19.42	14.87
9.17	7.70	19.85	15.28
9.28	7.79	19.97	15.39
9.73	8.03	20.37	15.77
10.03	8.13	20.69	16.07
10.35	8.34	20.99	16.36
10.71	8.51	21.33	16.59
10.88	8.65	21.87	17.22
11.21	9.85	22.08	17.41
11.55	9.02	23.17	18.63
11.78	9.14	23.45	19.98
11.97	9.32	24.05	19.67
12.35	9.56	24.85	20.55
12.74	9.72	25.62	21.63
12.90	9.86	25.88	21.64
13.27	10.09	26.43	22.55
13.59	10.28	27.13	23.47
14.18	10.73	27.87	24.46
14.45	10.88	28.54	25.46
14.83	11.13	29.14	26.31
14.92	11.08	30.68	28.72
15.21	11.44	31.17	29.51
15.71	11.71	31.87	30.70
16.05	12.24	32.65	32.04
16.41	12.31	33.45	33.47
17.82	12.61	33.87	34.27
17.11	12.82	34.38	35.29
17.42	13.08	34.92	36.31
17.92	13.57	35.31	37.12
	12.89%		
16.05	12.92	21.67	18.17
16.36	13.11	22.64	19.16
16.68	13.32	23.84	20.66
17.00	13.64	24.27	21.17
17.41	13.99	25.01	21.22
18.16	14.63	26.05	23.62
18.25	14.71	27.30	25.46
18.40	14.85	28.27	26.92
19.01	15.45	29.25	28.66
19.33	15.71	32.03	33.31
20.65	16.84	33.06	35.21
20.92	17.37	34.30	37.91

12.80%			
7.81	7.51	10.84	9.18
7.99	7.58	11.12	9.34
8.15	7.68	11.69	9.70
8.37	7.79	11.81	9.75
8.49	7.84	11.94	9.89
8.75	7.99	12.24	10.01
8.77	8.00	12.32	10.12
8.95	8.09	12.57	10.28
9.14	8.21	13.32	10.44
9.38	8.31	13.56	10.78
9.51	8.39	13.81	11.12
9.69	8.49	14.18	11.42
9.89	8.61	14.45	11.57
10.09	8.66	14.90	11.95
10.21	8.79	15.25	12.28
10.42	8.91	15.35	13.15
10.60	9.04		

Tammann, 1885

t	p				
	0%	13.19%	20.84%	34.41%	47.67%
31.66	35.0	33.7	32.4	29.6	25.8
39.33	53.4	51.1	48.4	44.3	39.0
44.53	70.2	67.0	63.8	58.0	51.2
55.11	118.7	112.5	108.3	98.0	86.7
58.85	141.7	134.6	129.0	117.5	103.8
63.54	175.7	166.6	159.6	145.9	128.7
68.59	219.9	208.1	199.2	182.0	149.8
71.92	254.0	240.1	230.7	209.9	187.3
76.18	303.9	287.3	275.1	250.3	219.8
80.85	367.9	347.9	333.4	303.7	267.1
82.41	391.7	370.5	355.9	323.2	283.4
86.97	468.4	442.9	424.7	385.5	338.3
90.15	529.1	500.0	480.0	436.0	384.3
92.60	580.1	548.1	525.2	477.6	418.9
95.35	642.3	606.5	581.6	528.0	464.0
100.27	767.3	735.91	694.9	632.3	556.3

%	p	%	p
100°			
5.80	744.2	37.11	608.2
9.93	729.8	39.25	598.5
18.37	699.0	43.12	577.4
27.07	661.6	46.93	555.0
31.75	639.8	53.98	504.6

Nicol, 1886

mol%	Dp					
	70°	75°	80°	85°	90°	95°
1.95	-8.5	-10.9	-13.3	-16.2	-19.7	-23.8
3.95	16.1	19.6	24.2	29.9	36.0	44.1
4.8	19.8	24.6	30.7	37.4	45.6	55.8
5.7	22.9	28.6	36.1	44.1	53.9	65.5
7.4	30.7	37.9	46.9	57.4	70.0	85.6
9.1	36.7	45.8	56.7	69.7	84.9	103.3
13.04	51.7	64.2	79.5	97.5	119.3	144.7
16.73	65.8	80.9	99.4	122.7	149.4	181.2
20.0	76.8	94.8	117.4	144.3	176.2	212.9

Emden, 1887

t	p	t	p
9.08%		29.45%	
19.40	16.1	20.50	15.5
24.40	21.75	25.26	20.55
30.31	31.0	30.00	27.0
36.20	42.95	34.29	34.65
40.72	54.55	35.32	36.65
45.70	70.5	40.34	47.9
50.74	91.0	45.48	62.2
55.74	116.45	49.63	77.15
60.61	145.9	55.60	103.2
65.21	180.05	59.91	126.5
69.69	219.55	64.40	155.0
76.19	289.45	69.81	196.6
80.21	341.35	77.29	271.2
		80.79	312.0
15.84	13.0	85.30	373.4
20.38	17.15	90.86	463.4
27.76	26.75	95.03	542.6
30.60	31.25		
34.78	39.6	37.54%	
41.33	56.55		
46.19	72.6	19.58	13.6
51.37	94.15	25.74	19.65
69.60	218.65	31.29	27.2
84.55	406.95	35.59	35.1
90.41	512.4	40.41	45.3
95.66	622.1	45.93	60.1
		49.88	73.3
16.44°		54.72	92.9
		59.47	116.2
22.22	18.4	65.08	149.6
25.70	22.75	71.11	195.6
30.50	30.25	75.51	235.65
35.10	39.2	80.40	287.3
40.60	52.65	85.68	354.9
46.38	70.1	90.51	427.5
49.95	85.0	95.36	511.7
55.25	110.0		
59.68	135.8	44.40%	
65.70	178.15	15.61	9.9
69.40	209.55	20.23	13.4
74.79	264.25	25.78	18.8
80.10	329.0	31.73	26.5
85.47	407.75	32.30	27.6
90.23	491.2	40.93	43.9
95.24	592.9	45.83	56.55
44.40%		44.40%	
51.63	75.5	76.13	225.9
55.31	88.9	80.72	272.1
60.52	114.4	85.54	331.25
66.78	152.65	90.18	395.8
71.52	186.25	93.87	455.5

Nicol, 1887 and 1888

t	p
sat.sol.	
65	120.9
75	179.5
85	256.4
95	362.8

Dieterici, 1891

%	p	%	p
	0°		
0	4.620	25.35	4.146
7.83	.483	36.29	3.953
14.52	.363	40.45	3.749

Lincoln and Klein, 1907

%	p	%	p
	25°		
0	23.76	33.12	20.05
10.46	22.78	43.07	18.54
24.51	21.21		

Speranski, 1909 and 1910

t	p	t	p
	sat.sol.		
23.8	16.27	34.8	29.57
25.8	18.35	35.8	31.06
27.8	20.41	36.8	32.81
29.8	22.67	37.8	34.57
31.8	25.25		

Applebey and Hughes, 1915

t	p	t	p
	sat.sol.		
119.88	744.4	121.20	772.3
120.40	755.5	121.23	775.0
120.49	757.7	122.07	791.0
120.58	758.7	122.56	800.4
120.77	763.45	122.89	806.3
121.18	772.2	123.17	811.6

Edgar and Swan, 1922

t	p	t	p
	sat.sol.		
20.09	13.30	26.56	19.31
20.96	13.92	27.65	20.58
23.03	15.84	28.99	22.08
23.35	16.04	29.81	23.26
25.39	18.01	29.84	23.20

Mondain-Monval, 1923

%	t	p
46.83	21	14.02
46.97	22	14.87
47.02	23	15.76
47.28	24	16.70
47.97	25	17.68

Adams and Merz, 1929

t	p	t	p
	sat.sol.		
10	7.13	30	23.07
15	9.85	40	38.81
20	13.53	50	62.21
25	17.73		

Dingemans and Dykgraef, 1948

t	p	%	p	%	p
		sat.sol.			
10.00	7.1	57.50	87.2	105.00	490.1
12.50	8.3	60.00	97.2	107.50	527.7
15.00	9.6	62.50	108.0	110.00	567.8
17.50	11.2	65.00	119.7	112.50	609.8
20.00	13.0	67.50	132.5	115.00	654.0
22.50	15.1	70.00	146.4	117.50	700.5
25.00	17.4	72.50	161.4	120.00	749.2
27.50	20.0	75.00	177.6	122.50	800
30.00	22.9	77.50	195.1	125.00	854
32.50	26.2	80.00	213.9	127.50	909
35.00	29.9	82.50	234.2	130.00	968
37.50	34.0	85.00	255.9	132.50	1028
40.00	38.6	87.50	279.1	137.50	1157
42.50	43.6	90.00	303.9	140.00	1225
45.00	49.2	92.50	330.3	142.50	1295
47.50	55.5	95.00	358.6	145.00	1368
50.00	62.3	97.50	388.6	147.50	1443
52.50	69.9	100.00	420.5	150.00	1520
55.00	78.2	102.50	454.3		

Legrand, 1835

%	b.t.	%	b.t.
8.51	101	54.61	112
15.75	102	56.75	113
22.00	103	58.74	114
27.49	104	60.59	115
32.29	105	62.29	116
36.55	106	63.87	117
40.37	107	65.34	118
43.79	108	66.73	119
46.89	109	68.00	120
49.70	110	69.21	121 sat.sol.
52.26	111		

Kremers, 1856

Sat. sol. b.t. = 122°

Berkeley and Applebey, 1911

Sat. sol. b.t. = 120.20°

Gerlach, 1886 and 1887

%	b.t.	%	b.t.
0	100	52.49	111
8.26	101	54.85	112
15.61	102	57.06	113
21.88	103	59.09	114
27.54	104	60.93	115
32.43	105	62.75	116
36.71	106	64.41	117
40.48	107	65.98	118
43.98	108	67.48	119
47.09	109	68.95	120
49.88	110		

Buchanan, 1899

%	b.t.	%	b.t.
620.3 mm(sat.sol.)		552.43 mm	
65.70	112.344	66.02	108.09
65.82	112.011	65.92	107.48
61.18	109.376	61.44	105.39
57.58	107.40	58.25	103.74
53.19	105.42	50.49	102.09
50.56	104.43		
47.23	103.18	63.46	108.19
43.09	101.91	63.01	107.59
39.44	100.89	58.01	105.22
35.55	99.90	52.24	102.69
30.78	98.88	50.38	101.55
27.51	98.24	48.16	100.72
		45.58	99.73
		42.57	98.80
760.0 mm		760 mm	
27.18	103.95	8.93	100.91
22.21	102.93	8.25	100.81
19.21	102.43	7.29	100.71
16.78	101.92	6.52	100.61
14.80	101.72	5.70	100.51
13.49	101.52	4.85	100.41
12.07	101.32	3.85	100.31
10.46	101.11		

Gerlach, 1926

%	b.t.	%	b.t.
1.2	100.1	26.2	103.8
2.3	100.3	32.1	104.9
4.5	100.6	39.6	106.8
8.6	101.1	46.9	108.9
14.4	101.8	46.6	110.5
15.9	102.1		

Freezing curve.

Kopp, 1840

%	f.t.
46.52	17.8
47.01	19.3

Poggiale, 1843

%	f.t.	%	f.t.
40.76	-6	52.63	50
44.37	0	54.54	60
45.74	10	56.47	70
46.70	16	58.45	80
47.24	20	60.58	90
48.82	30	62.71	100
50.55	40	69.26	120

Rüdorff, 1861

%	f.t.	%	f.t.
0.99	-0.4	9.09	-3.6
1.96	-0.75	10.71	-4.35
3.85	-1.5	12.28	-4.9
5.66	-2.35	13.79	-5.65
7.41	-2.9		

Mulder, 1866

%	f.t.	%	f.t.
42.17	0	51.81	40.25
42.89	4.25	52.89	51
43.79	9.5	59.39	79
45.11	13.5	62.19	91
45.80	16	64.26	99
46.21	17.25	68.39	117.5 b.t.
47.59	26		

Nordenskjöld, 1869

%	f.t.	%	f.t.
42.20	0.0	55.66	60.65
44.93	13.9	63.44	99.9
52.49	44.65	67.88	119.7

Rüdorff, 1872

%	f.t.	%	f.t.
5.66	-2.20	13.79	-5.55
7.41	-2.85	16.67	-6.60
9.09	-3.55	18.03	-7.20
10.71	-4.25	28.57	-12.30
12.28	-4.85		

de Coppet, 1872

%	f.t.	%	f.t.
10.71	-4.25	33.33	-15.5
16.67	-6.7	37.50	-18.7
23.08	-9.7	41.52	-21.67
28.57	-12.7		

Ditte, 1875

%	f.t.	%	f.t.
40.12	-15.7 E	44.63	15
40.01	0 (14+1)	45.54	18
41.51	2	46.16	21
41.53	4	47.46	26
43.07	8	48.17	29
43.28	10	49.85	36
44.13	13	53.19	51
		55.57	68

Guthrie, 1876

%	f.t.	%	f.t.
5	-2	30	-13
10	-4.2	35	-15.5
15	-6.3	40.80	-17.5 E
20	-8.4	42.34	0
25	-10.8		

Bodländer, 1891

%	f.t.
45.01	13
45.25	16.5
45.47	15

Etard, 1894

%	f.t.	%	f.t.
66.8	120	83.5	220
67.5	130	89.5	250
77.1	172	91.5	255
78.1	180	97.5	290
82.0	199	100.0	313

Jones, 1902

M	f.t.	M	f.t.
0.05	-0.198	1.0	-3.205
0.1	-0.364	1.5	-4.670
0.2	-0.706	2.0	-6.136
0.3	-1.040	2.5	-7.468
0.4	-1.348	3.0	-8.884
0.5	-1.672		

Jones and Getman, 1904

M	f.t.
1.00	-3.198
1.50	-4.669
2.00	-6.147
2.50	-7.468
3.00	-8.909

Massink, 1913

%	f.t.
44.6	10
46.8	20
47.47	25
49.31	35

Rodebush, 1918

%	f.t.
17.24	- 6.67
22.34	- 8.78
25.45	-10.17
30.82	-12.85
34.70	-15.08
38.46	-17.46 E

Mondain-Monval, 1923

%	f.t.
45.90	16.2
46.75	20.0
47.92	25.0

Foote, 1925

%	f.t.
50.22	35
48.10	25

Findlay and Cruickshank, 1926

46.77 % f.t. = 20°

Kordes, 1926

mol%	f.t.
0	0
11	-18 E
100	308

Chrétien, 1929

%	f.t.	%	f.t.
42.3	0	63.7	100
42.9	25	67.4	118
53.3	50	sat.sol.	b.t.120.1
58.6	75		

E : -17.7°

Foote and Vance, 1929

%	f.t.
42.13	0
43.99	8
48.04	25
50.15	35

Nikitina, 1933

%	f.t.	%	f.t.
42.40	0	52.30	50
43.28	5	54.13	52
45.70	15	55.47	60
46.80	20	58.84	76
47.17	21	59.90	80
47.85	25	60.00	81
51.30	40	61.72	90
52.06	44	63.04	98
		63.70	100

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
64.1	102	86.6	218
66.6	118	88.3	227
68.1	131	90.5	241
72.6	143	92.2	251
76.2	163	94.6	265
77.2	167	96.0	275
80.9	187	97.6	291
82.7	198	100.0	308
85.5	211		

Khitrova, 1954

%	f.t.	%	f.t.
5	-2.7	23	-3.4
15	-5.4	32	-12.4
18	-6.3		

Robinson , 1935

Isopiestic solutions

m_1	m_2	m_1	m_2	m_1	m_2
25°					
0.1040	0.1052	1.419	1.529	3.257	3.840
.2036	.2059	.456	.569	.272	3.863
.2052	.2077	.543	.674	.391	4.038
.3365	.3431	.647	.798	.434	.087
.4307	.4428	.687	.845	.585	.315
.5090	.5236	.868	2.060	.698	.467
.5732	.5935	2.304	.596	.888	.735
.7063	.7357	.418	.738	4.076	5.000
.8404	.8774	.482	.814	.095	.016
.8724	.9159	.560	.912	.134	.072
1.109	1.175	.595	.956	.228	.194
.111	.181	.731	3.127	.431	.505
.132	.205	.734	.138	.516	.628
.211	.293	.809	.237	.542	.671
.235	.317	3.000	.506	.616	.759
.342	.441	.130	.672	.81	6.025
		.167	.719		

m_1= Potassium chloride m_2= Sodium nitrate

Robinson and Stokes, 1949

m	osmotic coefficient
	25°
0.1	0.921
.2	.902
.3	.890
.4	.881
.5	.873
.6	.867
.7	.862
.8	.858
.9	.854
1.0	.851
.2	.845
.4	.839
.6	.835
.8	.830
2.0	.826
2.5	.817
3.0	.810
3.5	.804
4.0	.797
4.5	.792
5.0	.788
5.5	.787
6.0	.788

Properties of phases .

Grassi, 1851

%	t	d
27.529	20.9	1.2026

Kremers, 1855 and 1861

%	d	%	d
	19.5°		
12.057	1.0825	39.860	1.3150
22.736	.1646	46.251	.3781
31.987	.2426		

t	d			
% interpolated from d 19.5°				
0	1.1125	1.2022	1.2747	1.3472
19.5	.1044	.1910	.2619	.3332
40	.0939	.1780	.2475	.3177
60	.0819	.1645	.2327	.3023
80	.0687	.1499	.2174	.2863
100	.0544	.1346	.2015	.2700

Schiff, 1858 and 1859

%	d	%	d
	20.2°		
1	1.0047	26	1.1883
2	.0113	27	.1965
3	.0129	28	.2048
4	.0245	29	.2132
5	.0313	30	.2217
6	.0380	31	.2303
7	.0449	32	.2390
8	.0518	33	.2478
9	.0587	34	.2566
10	.0657	35	.2656
11	.0727	36	.2747
12	.0798	37	.2840
13	.0879	38	.2934
14	.0942	39	.3031
15	.1015	40	.3131
16	.1089	41	.3231
17	.1164	42	.3331
18	.1240	43	.3432
19	.1317	44	.3533
20	.1397	45	.3635
21	.1477	46	.3736
22	.1557	47	.3839
23	.1638	48	.3943
24	.1719	49	.4049
25	.1801	50	.4155

Fouqué, 1867

%	t	d	
		0°	t
0	9	-	1.1439
1.0	22	1.0071	1.0050
3.58	22	1.0252	1.0221
15.97	20	1.1507	1.1389

Thomsen, 1871

mol%	d	mol%	d
	18°		
0	0.9986	2	1.0600
0.5	1.0160	3.8	1.1137
1	1.0311	9	1.2474

van der Willigen, 1874

%	d	%	d
	22.80°		
16.86	1.11778	39.14	1.30618
26.28	.19274	44.35	.35774
33.89	.25785		

Forster, 1878

%	d
25.0°	
0	0.9971
21.23	1.1519
42.90	1.3442

Kohlbrausch, 1879

%	d	%	d
18°			
5	1.0327	20	1.1435
10	1.0681	30	1.2278

Rönnberg, 1880

t	d	t	d
5 %			
3.5	1.03342	21.5	1.02965
9.0	.03242	25.7	.02792
14.5	.03121		
10 %			
3.5	1.06480	20.7	1.06011
9.0	.06353	25.7	.05820
14.5	.06205		
15 %			
3.5	1.09445	21.0	1.08835
9.5	.09264	25.7	.08621
14.5	.09091		
20 %			
3.5	1.12239	21.0	1.11570
9.5	.12035	25.7	.11342
15.0	.11823		
25 %			
3.5	1.14904	20.7	1.14136
9.5	.14668	26.0	.13841
15.0	.14421		

Volkmann, 1882

%	d	%	d
12°		15-16°	
16.80	1.1311	7.35	1.0490
30.43	.2301	15.37	.1067
38.33	.3022	31.02	.2338
		44.49	.3623

Nicol, 1883

mol%	d 20°	d 40°
0	0.99823	0.99224
1.96	1.05793	1.04917
2.44	.07176	.06255
4.76	.13611	.12479

Kanonnikoff, 1885

%	t	d
0	20.6	0.9981
18.04	20.6	1.1269

Röntgers and Schneider, 1886

%	t	d
5.70	18.1	1.0383
11.41	17.9	1.0791

Nicol, 1887 and 1888

t	d
5.5 N (20°)	
20	1.4547
40	1.12550
98	1.0483

Bodländer, 1891

%	d	%	d
16.5°		15°	
45.01	1.3688	45.47	1.3720
45.25	1.3733		

Gilbault, 1897

%	d	%	d
20°			
0	0.9982	30	1.2217
5	1.0314	40	.3132
10	.0657	50.07	.4162
20	.1398		

Sentis, 1897

mol%	t	d	mol%	t	d
10	2.6	1.2752	5	11.6	1.1459
10	12.2	.2683	4	19.3	.1144
10	20.4	.2621	3	16.8	.0881
10	23.8	.2612	2	16.5	.0594
8	18.0	.2168	1	16.5	.0278
6	18.4	.1664	1	11.9	.0312

Barnes and Scott, 1898

%	d	%	d
	20.1°		
0	0.9982	17.370	1.1228
1.589	1.0096	23.24	.1696
4.241	.0273	31.72	.2407
7.039	.0468	35.65	.2765
9.665	.0656	42.05	.3380
11.915	.0819		

Pann, 1901

%	d		
	10°	18°	30°
0	0.999727	0.998622	0.995673
5.39	1.03731	1.03545	-
8.35	.05837	.05637	-
20.24	.15108	.14630	1.13934
35.57	.27962	.27461	.26690
50.87	-	-	.38991

Berkeley, 1904

%	t	d
	sat.sol.	
42.30	0.30	1.3530
45.80	15.45	1.3769
49.17	30.00	1.3992
52.16	44.50	1.4210
55.46	60.00	1.4446
58.87	76.15	1.4701
61.77	90.25	1.4920
67.09	119.0	1.5374

Loewenfeld, 1905

%	d
15°	
0	0.9993
9.4	1.0647
17.2	.1232
23.1	.1705
28.7	.2176

Cheneveau, 1907

%	d	%	d
	15°		
40.09	1.3191	22.82	1.1680
37.00	.2910	18.77	.1356
33.71	.2594	14.48	.1023
30.25	.2293	9.96	.0692
26.64	.1994	5.13	.0343
		0	0.9991

Getman, 1907

%	d	%	d
	18°		
0	0.9986	24.76	1.2113
9.37	1.0665	32.09	1.2465
17.75	1.1271		

Getman and Wilson, 1908

%	t	d
16.9	22.8	1.1178
44.4	22.8	.3580

Rubien, 1911

M	d	M	d
	18°		
0	1.33327	1.025	1.34274
0.1034	.33430	2.032	.35111
0.2058	.33528	3.979	.36606
0.5116	.33813		

Chernai, 1912

%	t	d	%	t	d
		sat.sol.			
59.68	80	1.477	49.03	30	1.406
57.63	70	.467	46.81	20	.387
55.48	60	.456	44.60	10	.377
53.27	50	.437	42.20	0	.358
51.20	40	.418	40.48	-10	.342

Buchanan, 1912-13

m	d 15.0°	d 19.5°
0	0.995126	0.998332
0.008	0.999604	-
0.016	1.000036	0.999271
0.031	1.000887	1.000171
0.062	1.002623	.001954
0.125	-	.005479
0.25	-	.012472
0.5	-	.026137
1	-	1.0525
2	-	.1012
3	-	.1459
4	-	.1871
5	-	.2247
6	-	.2592
7	-	.2918
8	-	.3231
9	-	.3517

Fedotiev and Koltunov, 1914

%	t	d
	sat.sol.	
42.31	0	1.354
45.62	15	.375
49.01	30	.401

Cornec and Chrétien, 1925

%	t	d
	sat.sol.	
42.3	0	1.352
47.8	25	.392
53.2	50	.427
58.7	75	.469
63.7	100	.507

Manchot, Jahrstorfer and Zepter, 1924

%	d	%	d
	25°		
9.181	1.0560	35.704	1.2152
11.136	.0677	18.362	.1106
18.447	.1141	36.724	.2189
25.588	.1543		

Gerlach, 1926

%	t	d	%	t	d
1.2	100.1	1.005	26.2	103.8	1.188
2.3	100.3	.016	32.1	104.9	.247
4.5	100.6	.028	39.6	106.8	.313
8.6	101.1	.060	46.9	108.9	.382
14.4	101.8	.096	46.6	110.5	.390
15.9	102.1	.108			

Hrynakowski, 1927

t	d
sat.sol.	
50.20	1.4383
49.85	1.4415

Geffcken, 1929

m	d
25°	
0	0.99707
1.7629	1.08754
4.6610	1.20835
6.6176	1.27547
9.1926	1.35045

Chrétien, 1929 and Cornec and Neumeister, 1929

%	t	d
	sat.sol.	
42.3	0	1.352
47.8	25	.392
53.3	50	.427
58.6	75	.469
63.7	100	.507
67.4	118	.535

Spacu and Popper, 1934

%	d	%	d
	20°		
0	0.99823	16.960	1.11879
8.489	1.05643	25.591	1.18678
16.960	1.11858		

Guillaume, 1946

%	d
20°	
8.58	1.0574
17.7	.1260
22.44	.1624
29.5	.2211
39.6	.3142

Nicol, 1886 and 1887

mol%	molar volume	mol%	molar volume
20°			
1.95	1858.22	7.4	2051.98
3.85	1921.31	9.1	2125.07
5.65	1987.77	10.7	2196.35

t	rel.vol.	t	rel.vol.
1.95 mol%		7.4 mol%	
20	100.000	20	100.000
45.6	101.088	45.5	101.396
50.4	101.325	50.1	101.665
56.1	101.633	56.4	102.047
61.6	101.936	61.2	102.351
67.1	102.277	67.1	102.725
72.2	102.593	72.3	103.068
78.5	102.999	78.5	103.486
3.85 mol%		9.1 mol	
20	100.000	20	100.000
45.4	101.175	45.6	101.435
50.5	101.457	50.7	101.739
56.2	101.786	56.6	102.102
61.3	102.091	61.6	102.425
67.2	102.455	67.9	102.823
72.4	102.796	72.1	103.140
78.5	103.191	78.0	103.480
5.65 mol%		10.7 mol%	
20	100.000	20	100.000
45.5	101.333	45.6	101.456
50.1	101.597	50.7	101.755
56.4	101.970	56.6	102.119
61.2	102.276	61.6	102.446
67.1	102.639	67.9	102.842
72.3	102.979	72.7	103.156
78.5	103.395	78.0	103.499

Pohl, 1906

%	relative volume
13.9°	
0	1.000
7.14	0.933
14.90	.792
20.91	.751
26.09	.702
32.60	.610

de Lannoy, 1895

t	relative vol. 4%	10%	20%	40%
0	1.00000	1.00000	1.00000	1.00000
10	.00150	.00265	.00409	.00508
20	.00379	.00610	.00852	.01020
30	.00690	.01012	.01338	.01600
40	.01080	.01478	.01870	.02200
50	.01546	.01958	.02435	.02802
60	.02066	.02510	.03040	.03420
70	.02652	.03124	.03660	.04060
80	.03298	.03762	.04334	.04720
90	.04026	.04458	-	-

Grassi, 1851

%	t	π
27.529	20.9	29.5

Röntgers and Schneider, 1886

%	t	π (apparent)
11.41	17.87	0.851
10.38	17.52	0.868
5.70	17.79	0:920

Gilbault, 1897

%	π	%	π
20°			
0	44.37	30	31.18
5	41.87	40	27.59
10	39.18	50.07	24.30
20	35.10		

Mikhailov, 1956 (fig.)

molality	Sound Velocity m/sec 0°	20°	40°	60°	80°
0.1	1435	1490	1530	1555	1550
1.0	1485	1530	1565	1585	1570
2.0	1540	1570	1590	1600	1590
5.0	1655	1675	1680	1670	1650
7.5	1745	1748	1740	1725	1705
10.0	1820	1805	1790	1765	1740

Viscosity and surface tension

Getman, 1907

%	η
18°	
0	1036
9.37	1146
17.75	1299
24.76	1549
32.09	1789

Volkmann, 1882

%	σ
12°	
16.80	76.9
30.43	78.9
38.33	80.4
15-16°	
7.35	74.8
15.37	76.0
31.02	78.8
44.49	82.7

Sentis, 1897

mol %	t	σ
0	25.1	72.3
0	13.5	74.0
1	11.9	75.4
1	16.5	74.2
2	16.5	75.0
3	16.8	75.7
4	19.3	75.5
5	11.6	77.4
6	18.4	76.9
8	18.0	77.9
10	2.6	82.1
10	12.2	80.4
10	20.4	79.5
10	23.8	79.2

Forch, 1899

molarity(18°)	t	σ	$\tau.10^5$
0.2237	22.37	-	196
0.2237	13.97	73.26	-
0.4477	20.88	-	203
0.4477	14.11	73.55	-
1.344	21.02	-	207
1.344	15.49	74.54	-
1.797	17.98	74.95	-
2.693	22.83	-	174
2.693	13.92	75.96	-
3.593	17.55	76.90	-
4.5	21.28	-	182
4.5	13.95	77.99	-

Pann, 1901

%	σ 10°	σ 30°
0	74.0096	71.0028
5.39	75.0288	-
8.35	75.3346	-
20.24	77.4434	74.6662
35.57	80.6736	78.2824
50.87	-	82.6532

Forch, 1905

molarity	t	σ
1.528	15.6	78.71
1.998	15.4	79.35
2.557	15.4	80.01
3.003	15.2	80.71
3.927	16.4	81.70
5.063	15.2	83.50

Loewenfeld, 1905

%	σ	%	σ
15°			
0	75.65	23.1	79.55
9.4	76.23	28.7	79.19
17.2	77.56		

Hrynakowski, 1927

t	σ
sat.sol.	
50.20	77.91

Optical and electrical properties

van der Willigen, 1874

spectral lines	n 16.86%	26.28%	33.89%	39.14%	44.35%
			22.80°		
A	1.34734	1.35809	1.36712	1.37351	1.37998
a	.34824	.35905	.36808	.37453	.38104
B	.34895	.35985	.36888	.37540	.38189
c	.34976	.36068	.36975	.37626	.38283
D	.35183	.36289	.37210	.37865	.38535
E	.35441	.36566	.37504	.38171	.38856
b	.35490	.36616	.37557	.38436	.39134
F	.35661	.36800	.37757	.38436	.39134
G	.35889	.37047	.38018	.38714	.39422
G	.36070	.37244	.38233	.38940	.39659
H	.36145	.37327	.38324	.39036	.39762
H	.36269	.37465	.38471	.39190	.39928
H	.36412	.37628	.38646	.39374	.40121

Forster, 1878

%	n_D
25.0°	
0	1.33246
21.23	1.35653
42.90	1.38339

Kanonnikoff, 1885

%	t	n H_α	D	H_β
0	20.6	1.33105	1.33297	1.33702
18.04	20.6	1.351375	1.353625	1.358606

Borgesius, 1895

%	t	n_D
0	20	1.33300
1.31	21.3	.33154
5.03	21.1	.33606
17.37	19.0	.35369
17.37	23.4	.35325

Cheneveau, 1907

%	n_D	%	n_D
	15°		
0	1.3334	26.64	1.3644
5.13	.3393	30.25	.3688
9.96	.3448	33.71	.3733
14.48	.3501	37.00	.3775
18.77	.3551	40.09	.3812
22.82	.3597		

Getman and Wilson, 1908

%	t	n_D
0	22.6	1.33309
16.9	22.8	1.35183
44.4	22.8	1.38538

Rubien, 1911

M	n_D	M	n_D
	18°		
0	1.33327	1.025	1.34274
0.1034	.33430	2.032	.35111
0.2058	.33528	3.979	.36606
0.5116	.33813		

Heydweiller, 1913

M	n_D	M	n_D
	18°		
0.1	1.33427	1.0	1.34250
0.2	1.33522	2.0	1.35086
0.5	1.33802	4.0	1.36622

Heydweiller, 1913

M	n_D	M	n_D
	18°		
0	1.33327	1.0	1.34313
0.1	.33431	2.0	.35229
0.2	.33532	4.0	.36916
0.5	.33833		

Geffcken, 1929

m	n_{He}	m	n_{He}
	25°		
0	1.332590	6.6176	1.374336
1.7629	1.347103	9.1926	1.384295
4.6610	1.365012		

Spacu and Popper, 1934

%	n_{He}	%	n_{He}
	20°		
0	1.3324865	16.889	1.371706
8.489	.342025	25.591	.361766
16.960	.351677		

Guillaume, 1946

%	d	n_{578}	* $\alpha^{mol}_{magn} \cdot 10^6$
		20°	
0	-	-	3.974
8.58	1.0574	1.3435	.738
17.7	.1260	.3538	.475
22.44	.1624	.3598	.341
29.5	.2211	.3679	.141
39.6	.3142	.3811	2.86

* in radians, gauss, centim.

Andersen and Asmussen, 1932

%	Verdet's constant	$(\alpha)^{mol}_{magn.}$
	0°	
43.26	0.0174	3.35
23.91	0.0164	3.10

Okazaki, 1933

%	Verdet's constant (3514 Å)
28°	
5.49	0.04149
12.35	.04107
20.89	.04088
29.14	.04023
38.68	.03927
43.46	.03916

Kohlrausch, 1879

%	κ
18°	
5	433
10	778
20	12.96
30	15.97

Jones and Getman, 1904

M	λ	M	λ
		0°	
0.5	44.90	1.0	34.16
0.1	40.41	1.50	30.96
0.2	36.90		

Sloan, 1910

M	λ	M	λ
		0°	
0.531	36.41	4.250	51.50
1.0625	43.20	8.50	54.72
2.125	47.81	17.00	57.41

Clausen, 1912

M	t	λ	$\tau \cdot 10^4$
0.5116	6	55.49	283.9
	18	73.65	376.8
	30	93.22	476.9
1.025	6	49.37	505.9
	18	65.30	669.3
	30	82.38	844.4
2.032	6	41.14	835.9
	18	54.24	1103
	30	68.37	1389
3.981	6	29.68	1182
	18	39.35	1567
	30	49.50	1971

de Malleman and Guillaume, 1945

22.44 %   d = 1.1624 $(\alpha)_{magn.} \cdot 10^5 = 9.76$

Heat constants.

Schüller, 1869

%	U	%	U
0	1	23.08	0.8341
9.09	0.9320	28.57	.7998
16.67	0.8768	33.33	.7673

Thomsen, 1871

mol%	U	mol%	U
	18°		
0.5	0.975	3.8	0.863
1	0.950	9	0.769
2	0.918		

Winkelmann, 1873

%	U	%	U
3.03	0.9707	19.19	0.8559
3.73	.9658	25.03	.8417
4.81	.9523	31.29	.8153
5.62	.9442	40.06	.7820
8.40	.9234	49.98	.7576
11.36	.9025	57.97	.7376
16.64	.8700	70.09	.7121

Marignac, 1876

mol%	U (22-52°)
1	0.9576
2	.9220
3.8	.8712
7.4	.7946
11.8	.7299

Tendt, 1899

t	U	t	U	t	U
14.43%		24.64%		46.88%	
35.4	0.9594	33.6	0.8262	34.7	0.7081
38.5	.9341	34.8	.8271	42.5	.7112
39.3	.9354	38.6	.8230	43.7	.7129
45.7	.9301	43.6	.8292	46.5	.7153
43.2	.9266	47.5	.8269	47.4	.7143
49.6	.9406	52.3	.8335	50.1	.7180
53.1	.9498	54.2	.8364	50.5	.7185
				51.3	.7207

Randall and Rossini, 1929

m	U	m	U
	25°		
0.00	0.9979	0.20	0.9804
0.01	.9970	0.35	.9687
0.02	.9960	0.50	.9579
0.05	.9933	0.75	.9416
0.10	.9888	1.00	.9271

Winkelmann, 1873

%	Q diss (by g salt) 0°	50°
3.03	-62.32	-50.0
3.79	61.47	50.7
4.81	60.90	49.3
5.62	60.25	49.3
8.40	58.28	48.5
11.36	55.65	46.9
16.64	52.34	45.9
19.19	50.56	45.2
25.0	46.5	43.7
31.3	44.2	42.2
40.0	41.9	40.5
50.0	38.4	38.8
60.0	36.3	37.9
70.0	34.5	36.4

Scholz, 1892

M	Q diss (by g salt)
0.0625	-69.93
0.125	68.53
0.25	66.46
1.0	58.75
1.5	54.95
4.0	43.18

Vareli-Thevenet, 1902

%	Q diss (by g salt)	%	Q diss (by g salt)
0.2	-76.0	12.2	-41.7
0.4	65.7	13.8	39.8
0.6	62.8	15.2	38.1
0.8	59.1	16.6	34.6
0.9	58.8	23.0	32.5
1.9	57.5	28.5	26.9
3.8	54.0	33.3	21.5
5.6	49.5	37.5	20.7
7.4	47.4	41.1	17.7
9.0	45.8	42.2	11.1
10.7	43.2	42.5	2.0

Kempanski, 1916

mol%	Q diss (by mole salt) 30°	40°	50°	60°
12.5	-3100	-3131	-3148	-3173
13.3	3052	3087	3109	3130
14.3	3003	3044	3071	3096
15.4	2955	3000	3033	3063
16.7	2905	2957	2994	3029
18.2	-	2913	2956	2996
20	-	-	-	2962

Rümelin, 1907

%	Q dil	%	Q dil
4.48	-3.49	19.11	-44.7
4.51	-3.52	23.91	-58.7
8.57	-10.49	23.95	-59.7
8.63	-10.71	28.50	-71.4
13.56	-25.6	28.55	-71.6
13.60	-25.8	32.00	-85.4
15.83	-32.6	32.08	-85.7
15.83	-33	42.58	-104.4
19.08	-43.9	42.68	-104.8

Q dil (differentiel) by mole water .

Ståhlberg, 1914 - 1915

mol %	Q sol (integral)	mol %	Q sol (integral)
11.8	-1813	3.85	- 821
10.7	-1731	3.23	- 689
9.6	-1631	2.91	- 619
8.5	-1512	2.60	- 535
7.4	-1401	2.28	- 450
6.4	-1237	1.96	- 378
5.66	-1145	1.64	- 294
5.06	-1042	1.32	- 185
4.46	- 928	1.00	- 106

Mondain-Monval, 1923

sat.sol.	Q dil. 0°	16°
	-2570	-2020

Jäger, 1891

%	coefficient of thermic conductivity
0	100
20	94.9
22	94.1
40	92.7
44	90.4

Water (H_2O) + Sodium metaphosphate ($NaPO_3$)

Tammann, 1885

%	p	%	p
	100°		
8.45	749.6	17.15	736.2
12.76	743.5	26.24	717.9
13.96	742.1		

Morey, 1953 (fig.)

%	f.t.	
-	40.8	3 aq. - 2 aq.
70	147	
73.9	159	
-	169	(1+1) incongruent
76.5	210	
78	235	
79.4	256	
84.9	305	
-	343	(1+2) incongruent
92	348	
93	402	$NaPO_3$ II
-	443	tr.t.
96	517	$NaPO_3$ I
100	620	

Marignac, 1876

mol%	U
(24 - 55°)	
4	0.8495
2	0.9129
1	0.9525
0.5	0.9761

Water (H_2O) + Sodium borate ($NaBO_2$)

Poggiale, 1843

%	f.t.	%	f.t.
1.47	0	15.32	60
2.36	10	19.49	70
3.90	20	23.77	80
5.66	30	28.64	90
8.08	40	35.55	100
11.44	50		

Water (H_2O) + Sodium vanadate ($NaVO_3$)

Mc Adam and Pierle, 1912

%	f.t.		%	f.t.	
13.22	25		17.42	25	
23.04	40	+ 0 aq.	20.78	40	+ 2 aq.
34.22	50		24.79	60	
40.60	60		28.05	75	

Water (H_2O) + Disodium orthophosphite (Na_2HPO_3)

Italiener, 1917

%	f.t.	%	f.t.
80.72	0	84.96	29.9
81.16	20	91.08-92.80	38.0
81.44-82.98	25.2	93.68	42.7

Water (H_2O) + Sodium perchlorate ($NaClO_4$)

Freeth, 1924

%	f.t.
10	-0.3
20	-6.8
30	-11.1
40	-17.8
45	-22
56	-32 E
62.54	0 1 aq.
65.51	15
68.71	30
70.88	40
73.16	50
73.3	50.8 1aq.-0aq.
74.3	60 0 aq.
75.0	75

Cornec and Dickely, 1926

%	f.t.	%	f.t.
76.75	100	70.38	38
75.01	75	67.82	30
73.94	55	65.63	25
73.26 tr.t.	50	62.87	15

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.930	1.6	0.925
0.2	0.920	1.8	0.930
0.3	0.915	2.0	0.934
0.4	0.912	2.5	0.947
0.5	0.910	3.0	0.960
0.6	0.909	3.5	0.975
0.7	0.910	4.0	0.991
0.8	0.911	4.5	1.008
0.9	0.912	5.0	1.025
1.0	0.913	5.5	1.042
1.2	0.916	6.0	1.060
1.4	0.920		

Miller and Sheridan, 1956

Isopiestic solutions.

m_p	m_s	m_p	m_s
	25°		
6.14	3.88	15.14	7.65
10.85	5.81	16.25	8.11

m_p= molality of sodium perchlorate.

m_s= " " sulfuric acid.

Carlson, 1910

c	t	d
107.6	15	1.666
123.4 tr.t.	50 +1 aq.	1.731
141.4	143	1.789

Lübben, 1914

M	d
	18°
0.5004	1.03749
1.001	.07555
2.002	.15013
3.991	.29488

Heydweiller, 1921

M	d
	18°
0.5	1.0388
1	.0769
2	.1524
3	.2247
4	.2969

Mazzucchelli and Pro, 1926

%	d 15°	d 25°	%	d 15°	d 25°
4.00	1.0261	1.0293	11.97	1.0830	1.0788
5.92	1.0393	1.0362	16.03	1.1141	1.1091
8.06	1.0544	1.0509	20.06	1.1460	1.1405

Cornec and Dickely, 1926 and 1927

%	t	d	%	t	d
		sat.sol.			
76.75	100	1.758	70.38	38	1.713
75.01	75	.757	67.82	25	.683
73.94	55	.756	65.63	15	.663
73.26	50	.749	62.87	0	-

Hantzsch and Düringen, 1928

%	d
20°	
0	0.99540
4.3442	1.0221
6.3596	.03473
20.034	.1285

Kohner, 1928

m	d	m	d
25°			
0	0.99107	4.00094	1.25270
1.00020	1.07133	5.00370	.30254
2.00248	1.13820	6.00284	.34810
2.77662	1.18532	7.00352	.39011

Hölemann and Kohner, 1931

%	d	%	d
25°			
0	0.99707	50.19	1.43730
24.35	1.17624	57.58	1.53167
41.94	1.34369		
35°			
0	0.99406	42.00	1.33566
24.38	1.17022	57.56	1.52169
45°			
0	0.99024	42.03	1.32713
24.33	1.16282	46.97	1.38092
39.07	1.29668	57.55	1.51170

Miller and Doran, 1956

0°: $d=0.9999 + 0.08749\,N - 0.0422\,N^{3/2}$ $\sigma=0.0007$

29.87°: $d=0.9957 + 0.7919\,N - 0.0185\,N^{3/2}$ $\sigma=0.00055$

49.3°: $d=0.9884 + 0.07558N - 0.000883N^{3/2}$ $\sigma=0.00035$

Lübben, 1914

M	$n^{18°}$ 4600.9 Å	3611.9	3467.0	3403.6
-	1.33878	1.34755	1.34961	1.35061
0.5004	.34274	.35142	.35349	.35449
1.001	.34628	.35511	.35718	.35818
2.002	.35319	.36207	.36416	.36517
3.991	.36578	.37477	.37687	.37687

M	3333.7 Å	3255.8	3133.3	3085.1
-	1.35572	1.35681	1.35944	1.36224
0.5004	.35961	.36070	.36334	.36615
1.001	.36331	.36440	.36704	.36986
2.002	.37032	.37142	.37406	.37690
3.991	.38307	.38418	.38684	.38971

M	2981.1 Å	2881.1	2748.7	2573.2	2312.9
0	1.35944	1.36224	1.36653	1.37359	1.38893
0.5004	.36334	.36615	.37044	.37753	.39291
1.001	.36704	.36986	.37416	.38129	.39668
2.002	.37406	.37690	.38122	.38837	.40379
3.991	.38684	.38971	.38405	.40129	.41679

Mazzucchelli and Vercillo, 1925

%	n_D 15°	25°
4.04	1.33628	1.33523
5.98	.33771	.33648
8.04	.33915	.33781
11.96	.34202	.34054
16.04	.34503	.34335
20.01	.34800	.34021

Hantzsch and Düringen,1928

%	n_D
20°	
0	1.33300
4.3442	.33579
6.3596	.33707
20.034	.34624

Hölemann and Kohner, 1931

%	n_D	%	n_D	%	n_D
25.0°		35.0°		45.0°	
0	1.33266	0	1.33142	0	1.32995
24.35	.34953	24.35	.34743	24.33	.34522
41.94	.34369	42.00	.36157	39.07	.35656
50.19	.37948	57.56	.37702	42.03	.35917
57.58	.53167			46.97	.36367
				57.55	.37461

Kohner, 1928

m	n_D	m	n_D
	25°		
0	1.33253	4.00094	1.35613
1.00020	.33978	5.00370	.36040
2.00248	.34597	6.00284	.36422
2.77662	.35032	7.00352	.36772

Heydweiller, 1921

M	λ	M	λ
	18°		
0.5	71.7	3	46
1	65	4	38.76
2	55.05		

Miller, 1956 (fig.)

M		λ	
	0°	29.9°	50°
1	42	82	120
2	35	70	98
4	24	49	78
8	9	20	29
9.7	6	12	-

Water (H_2O) + Sodium permanganate ($NaMnO_4$)

White and Miller, 1953

%	f.t.	%	f.t.	
0	0	50	+2.25	
2	-0.5	52	6.6	
4	-0.9	54	10.8	
6	-1.4	56	15.0	
8	-2.0	58	18.9	
10	-2.5	60	22.5	
12	-3.2	62	25.8	
14	-3.8	64	30.7	
16	-4.6	66	31.3	
18	-5.3	68	33.7	
20	-6.1	70	35.3	
22	-7.0	72	36.0	(3+1)
24	-7.7	74	35.3	
26	-8.5	75	34.0	
28	-9.5	75.2	33.7	E
30	-10.3	76.0	36.4	
32	-11.2	78.0	45.6	
34	-12.2	80	51.5	
36	-13.2	82	59.0	
38	-14.1	83	62.5	
40	-15.0	84	65.2	
41.4	-15.8 E	86	67.7	
42	-14.7	88	68.7	(1+1)
44	-10.7	90	68.0	
46	-6.2	90.7	66.2	E
48	-2.0	92	66.5	
		94	67.6	
		96	67.6	

%		d		
	1.35°	13.00°	25.30°	37.45°
5.32	1.0404	1.0390	1.0366	1.0319
10.58	.0843	.0817	.0776	.0725
14.05	.1148	.1110	.1062	.1005
17.62	.1471	.1429	.1371	.1310
25.60	.2243	.2128	.2110	.2034

Water (H_2O) + Sodium phosphate (Na_3PO_4)

Tammann, 1885

%	p	%	p
	100°		
5.80	747.5	19.77	716.6
9.43	738.4	25.67	704.9
14.19	728.5	31.18	690.2

Lescoeur, 1890

t	p 15aq.	p Xaq.	t	p 15aq.	p Xaq.
0	1.8	3.1	30	18.0	26.5
5	2.7	4.3	35	26.0	36.0
10	3.9	5.8	40	38.5	f.t.
15	5.7	8.5	60	113	-
20	8.1	12.8	80	272	-
25	11.9	18.0	99	596	-

Legrand, 1835

%	f.t.	%	f.t.
0.0	0.0	43.31	4.0
9.91	0.5	45.71	4.5
17.35	1.0	47.78	5.0
23.67	1.5	49.59	5.5
28.98	2.0	51.22	6.0
33.46	2.5	52.69	6.5
37.27	3.0	52.95	6.6
40.52	3.5		

Apfel, 1911

molality	f.t.	
0.26	0	12aq.
0.75	25	
0.98	37	10aq.
1.02	40	
1.09	44	
1.38	50	8aq.
1.595	55	
1.84	65	
1.99	70	
2.14	75	

Greenish,

12.64 % f.t. = 15°

Kasankin, 1891

t	d	capillary rise (in mm)
	sat.sol.	
16-17	1.0490	25.85
32-34	1.1252	24.07

Water (H_2O) + Disodium phosphate (Na_2HPO_4)

Tammann, 1885

%	p	%	p
	100°		
6.99	747.2	25.61	711.0
14.57	731.8	32.75	693.2
17.56	725.7	34.88	686.0
20.95	719.5	44.12	648.2
23.52	713.9	45.69	639.0

Menzel and Sieg, 1932

t	p	t	p
	sat.sol. with 12aq.		
15.5	12.5	25.0	22.6
20.0	16.5	27.5	26.3

dissociation	t	p
12aq. - 7aq.	20	13.0
7aq. - 2aq.	20	10.0
2aq. - 0	20	2.6
	25	3.8
	30	5.7

Debray, 1868

t	p dissoc. 24 - 14aq.	p dissoc. 14 - 0aq.
12.3	7.4	4.8
16.3	9.9	6.9
20.7	14.1	9.4
24.9	18.2	12.9
31.5	30.2	21.3
36.4	39.5	30.5
40.0	50.0	41.2

Gerlach, 1886

%	b.t.	%	b.t.
0	100	37.46	103.5
7.92	100.5	40.62	104
14.67	101	43.47	104.5
20.50	101.5	46.04	105
25.59	102	48.38	105.5
30.02	102.5	50.50	106
33.95	103	52.49	106.5

Mulder, 1866

%	f.t.	%	f.t.
1.44	0	37.70-38.12	39
"	4.25	38.98	40
2.63	4.5	45.80	52.75
"	5	47.61	59
4.85	13	48.69	70
5.92	16.5	48.93	76
9.91	22.25	49.29	85
13.79	25.5	49.70	99
21.20	31	49.59	99.3
22.54	32	45.20	105
26.69	34.5	44.66	105.57
-	trans.	44.19	106.4 b.t.

Shiomi, 1909

%	f.t.	%	f.t.
1.71	0.65	40.70	45.14
3.43	10.26	43.37	47.23
3.46	10.36	44.45	48.23
4.97	15.11	44.49	48.33
7.31	20.24	44.55	50.22
10.73	25.15	44.87	55.17
10.96	25.40	44.94	55.27
11.05	25.50	45.35	60.23
17.22	30.21	46.84	70.26
17.75	30.26	48.66	80.39
18.97	30.76	50.71	89.74
23.60	33.04	51.76	94.75
23.89	33.14	51.70	95.86
31.21	36.27	51.22	96.86
32.20	37.27	50.29	99.57
35.43	40.29	50.52	99.77

hydrate	tr.t.
12aq. - 7aq.	36.45
7aq. - 2aq.	48.0
2aq. - 0aq.	95.2

D'Ans and Schreiner, 1910

%	f.t.	%	f.t.
2.44	0	44.13	48.35
10.34	15	47.72	59.7
31.56	35.40	48.75	71
35.40	40.3	49.53	91

Menzies and Humphrey, 1912

%	f.t.	
0	0	
0.54	-0.24	
1.40	-0.43	
-	-0.5	E
1.60	+0.05	12aq.
7.0	20	
10.7	25.0	
20.48	32.0	
25.0	34.0	
-	35.2	7aq.
34.15	39.2	
40.20	45.0	
-	48.3	
44.6	50.0	2aq.
45.30	60.0	
48.0	80.0	
50.5	90.2	
-	95	
51.0	96.2	0aq.
50.8	105.0	
49.7	120	

Hammick, Goadby and Booth, 1920

%	f.t.		
1.45	-0.47 E		
2.73	+6.00 12aq. α		
7.26	19.95		
8.93	22.77		
9.53	24.15		
10.90	25.75		
14.16	27.80		
15.87	28.65		
16.04	29.05		
17.18	29.50	tr.t.	12aq. α-β 29.6°
19.45	30.10		12 aq. α - 7 aq. 35.0°
20.08	30.90		
22.57	32.50		
24.63	33.70		
29.75	34.70		
31.15	36.50	12 aq. β 7aq.	
35.56	40.02		

Water (H_2O) + Monosodiumphosphate (NaH_2PO_4)

Tammann, 1885

%	p	%	p
	100°		
9.51	742.1	37.45	678.3
19.67	723.0	44.57	652.5
28.69	702.8	56.48	590.1

%	p	%	p
	100°		
9.12	745.2	53.54	611.8
20.32	723.1	63.34	538.9
47.96	639.3		

Robinson and Stokes, 1949

m	osmotic coefficient
25°	
0.1	0.911
.2	.884
.3	.864
.4	.847
.5	.832
.6	.819
.7	.808
.8	.798
.9	.789
1.0	.780
.2	.765
.4	.751
.6	.739
.8	.729
2.0	.721
2.5	.705
3.0	.696
3.5	.691
4.0	.691
4.5	.694
5.0	.699
5.5	.706
6.0	.713

Apfel, 1911

molality	%	f.t.	
3.01	36.1	0	
3.12	36.2	0	
3.82	45.8	18	
4.08	48.65	25	2aq.
4.03	48.55	25	
4.47	53.6	35.5	
4.70	56.4	40	
4.83	58.0	44	
5.06	60.7	44	laq. metast.
5.15	61.8	50	2aq.
5.32	63.8	55	laq. stable
5.46	65.5	58	
5.48	65.8	61	0aq.
5.49	65.9	65	
5.52	66.2	70	
5.60	67.2	75	
5.76	69.1	83	

Imadsu, 1912

%	f.t.	%	f.t.
36.65	0.10	58.16	40.20
37.14	1	58.47	40.55
38.07	3	58.76	41
38.95	5	58.98	42
41.14	10	59.71	45
43.41	15	61.33	50
46.01	20	62.10	52
48.62	25	63.08	55
49.04	27	63.39	56
50.41	28	63.65	57
51.55	30	63.93	58
52.15	31	64.21	60
53.47	33	64.45	62
53.94	34	64.91	65
54.64	35	65.55	69
55.89	37	67.46	80
58.00	40	69.26	90
		71.73	99.1

tr.t.

4aq. - 2aq. : 57.4% -20°

2aq. - laq. : 58.47% +40.8°

laq. - 0aq. +51.4°

Moore, 1895-96

M	d
18°	
0.00	0.9987
0.25	1.0184
0.5005	1.0391
1.001	1.0776
2.002	1.1677

Mason and Culvern, 1949

M	d	M	d
	25°		
0.001020	0.99716	0.1030	1.0065
.006859	0.99786	.1374	.0093
.008265	.99802	.1935	.0140
.01030	.99836	.2754	.0215
.01372	.99840	.3091	.0247
.02000	.99896	.4128	.0333
.02566	.99966	.5152	.0423
.02751	.99980	.6183	.0509
.03878	1.0007	.7783	.0642
.05152	.0019	.9274	.0867
.06864	.0035	1.367	.1129
.07778	.0038	2.061	.1676
.09274	.0060	3.885	.3053

Moore, 1895-96

M	η (water=1000)
18°	
0.00	1000
0.25	1076
0.5005	1182
1.001	1409
2.002	2313

Mason and Culvern, 1949

M	λ	M	λ
	25°		
0.001020	81.288	0.09274	64.779
.002064	80.471	.1030	64.077
.004007	78.308	.1374	61.937
.006859	76.632	.1935	59.291
.008265	76.368	.2754	55.810
.01030	75.782	.3091	54.883
.01372	74.399	.4128	51.732
.02000	73.324	.5152	49.281
.02566	72.202	.6183	47.031
.02751	70.978	.7783	43.868
.03878	69.957	.9274	41.504
.05152	68.411	1.367	35.493
.06864	66.709	2.061	28.115
.07778	65.757	3.885	15.513

Marignac, 1876

mol%	U
	24 - 55°
3.9	0.8444
2	.9070
1	.9499
0.5	.9704

Water (H_2O) + Halfsodium phosphate ($NaH_5P_2O_8$)

Parravano and Mieli, 1908

%	f.t.	%	f.t.
0	0	87.48	79.7
20.77	-5.7	88.65	85.0
26.92	-7.9	91.47	101.7
34.15	-11.4	92.67	104.5
56.66	-38	95.79	110.0
80.46	+34.0	95.86	110.7
81.82	41.0	97.99	119.0
83.68	51.7	100.00	126.5

Water (H_2O) + Sodium pyrophosphate ($Na_4P_2O_7$)

Tammann, 1885

%	p	%	p
	100°		
11.89	746.6	22.18	736.3
15.59	742.9	24.85	732.3
19.34	739.3		

Menzel and Sieg, 1932

t	p	t	p
	sat.sol. with 10aq.		
15.5	12.7	25.0	22.9
20.0	16.8	27.5	26.6

t	p	t	p
	dissociation 10aq. - 0aq.		
23	9.4	50.1	258.3
25.0	61.4	82.3	362.8

t	p	t	p
128	8.3	206	685
151	38.9	208	754
171	117	173	127
191	312	193	339
201	526		

%	f.t.	%	f.t.
2.14	-0.43	13.98	50.0
2.19	0.00	16.06	54.0
4.90	18.0	19.75	60.0
5.22	20.0	27.49	70.0
6.20	25.0	33.04	76.0
7.04	30.0	35.13	82.0
9.81	40.0	32.65	89.0
11.61	45.0	31.15	91.0
12.73	47.5		

Morey and Chen, 1956

t	p
374	216.5
400	272
500	461

Water (H_2O) + Trisodium pyrophosphate ($Na_3HP_2O_7$)

Hubbard, 1949

%		f.t.	%		f.t.
14.15	9aq.	20.7	27.50		32.9
14.28		21.0	32.70		36.1
20.00		26.5	22.15	1aq.	32.9
20.25		27.3	21.55		36.1
20.30		27.3	20.30		42.2
22.62		29.0	19.20		48.0
26.50		32.2	18.90		49.5

Water (H_2O) + Sodium arsenate (Na_3AsO_4)

Lescoeur, 1890

t	p dissoc. 15aq.	X aq.	sat.sol.
0	1	3.1	4.1
5	1.2	4.5	6.2
10	2.1	7.5	8.7
15	3.4	11.5	12.2
20	4.6	16.0	16.0
25	9.8	-	-
30	15.0	-	-
40	29.0	-	-
60	77.0	-	-
70	113.0	-	111.0
79	-	-	176.0
80	188.0	-	-
100	424.0	-	424.0

Schiff, 1860

%	d	%	d
	17°		
0	0.9988	5.89	1.0646
0.49	1.0041	6.38	.0703
0.98	.0095	6.87	.0760
1.47	.0149	7.36	.0817
1.96	.0203	7.85	.0874
2.45	.0258	8.34	.0932
2.94	.0313	8.83	.0990
3.43	.0368	9.32	.1048
3.92	.0423	9.81	.1107
4.41	.0478	10.30	.1166
4.91	.0534	10.79	.1225
5.40	.0590		
5.89	.0646		

Water (H_2O) + Disodium arsenate (Na_2HAsO_4)

Tammann, 1885

%	p	%	p
	100°		
8.06	746.4	27.20	708.7
14.83	734.4	31.40	698.4
18.28	728.8	32.95	693.7
22.79	719.6	40.72	666.9

Schiff, 1860

%	d	%	d
	14°		
0.46	1.0035	9.71	1.0944
0.93	.0077	10.17	.0993
1.38	.0119	10.64	.1043
1.85	.0161	11.10	.1093
2.31	.0205	11.56	.1144
2.78	.0249	12.03	.1195
3.24	.0292	12.49	.1246
3.90	.0336	12.95	.1296
4.16	.0581	13.42	.1348
4.63	.0426	13.88	.1400
5.09	.0471	14.34	.1452
5.56	.0517	14.80	.1505
6.02	.0563	15.27	.1557
6.18	.0610	15.79	.1611
6.44	.0656	16.19	.1665
7.51	.0703	16.66	.1718
7.87	.0750	17.12	.1773
8.33	.0798	17.58	.1805
8.79	.0846	18.04	.1882
9.25	.0895	18.50	.1938

Water (H_2O) + Monosodiumarsenate (NaH_2AsO_4)

Tammann, 1885

%	p	%	p
	100°		
9.26	745.9	37.07	692.5
18.50	732.6	41.69	678.4
28.33	713.6	48.65	653.0

Marignac, 1876

mol%	U
	26 - 57°
3.8	0.7884
2	.8707
1	.9264
0.5	.9595

Water (H_2O) + Sodium tetraborate ($Na_2B_4O_7$)

Tammann, 1915

%	p	%	p
	100°		
5.46	746.5	20.62	721.0
10.46	737.0	27.56	709.4
12.35	733.8	33.43	696.8
16.02	727.6		

Morey and Chen, 1956

t	P
V+L+C	
374	46.5
400	84
500	134
600	148

Gerlach, 1886

%	b.t.	%	b.t.
0	100	37.97	103
7.95	100.5	42.99	103.5
14.67	101	47.59	104
20.95	101.5	52.14	104.5
		52.89	104.6

Horn and van Wagener, 1903

%	f.t.	%	f.t.
1.28	5	13.79	57
1.57	10	16.60	60
2.71	21.5	17.08	61
3.76	30	17.21	62
5.30	37.5	18.03	65
7.50	45	19.61	70
9.50	50	23.90	80
11.73	54	28.98	90
12.43	55	34.34	100
13.04	56		

Blasdale and Slansky, 1939

%	f.t.	%	f.t.	% (4+1)
	(10+1)			
1.18	0	14.52	58.5	tr.t.1
1.44	5	-	60	14.82
1.76	10	16.65	60.8	tr.t.2
2.12	15	17.88	65	15.88
2.58	20	19.49	70	17.12
3.13	25	21.30	75	18.41
3.85	30	23.38	80	19.88
4.76	35	25.73	85	21.46
6.00	40	28.37	90	23.31
7.58	45	31.28	95	25.55
9.55	50	34.63	100	28.22
12.25	55	36.73	102.8	-

tr.t. 1 = (10+1) - (4+1)

tr.t. 2 = (10+1) - (5+1)

Keshan, 1955

%	f.t.	%	f.t.
	sat.sol.		
1.23	0	16.7	30
1.58	10	23.9	80
3.75	20	34.3	100

Benrath, 1942

%	f.t.	%	f.t.
38	108	57	130
43	116	60	132
48	121	63.1	135
52.8	125	66.2	137
53.9	126	69.1	140

Kasankin, 1891

t	d	capillary rise (in mm)
	sat.sol.	
16-17	1.0206	26.13
32.34	1.059	25.41

Dewar and Fleming, 1897

t	ε	t	ε
	15 %		
-202.0	39.4	-139.8	96.0
-195.3	67.2	-127.0	105.0
-188.4	79.5	-121.0	103.0
-176.0	86.3	-112.8	118.0
-152.7	90.5		

Water (H_2O) + Sodium pentaborate (NaB_5O_8)

Rollet and Chung-Ming, 1935

%	f.t.	%	f.t.
1.35	-0.53 ice	41.9	95
2.68	-0.91	45.5	90
3.94	-1.24	47.3	93
5.8	-1.70 E	49.0	96
6.0	0 5 aq.	50.9	99
8.6	+13.5	51.8	100 5aq. - 1aq.
10.7	20	52.7	103 1aq.
15.9	35	54.3	107
22.7	50	54.5	107.5
27.45	59.6	54.8	108.3
33.7	70.8	55.0	109
38.8	80	55.3	109.6
53.5	102 5aq.metast.	55.5	104.5 5aq.metast.

Blasdale and Slansky, 1939

%	f.t.	%	f.t.	%	f.t.
		5 aq.			
6.28	0	15.60	35	32.25	70
7.10	5	17.40	40	34.98	75
8.10	10	19.63	45	37.84	80
9.30	15	21.80	50	40.77	85
10.55	20	24.30	55	43.80	90
12.20	25	26.90	60	47.05	95
13.75	30	29.35	65	50.30	100

H_2O + Sodium thioantimonite (Na_3SbS_3)

Sklyarenko, Kaplan and Lurie, 1950

M	ϰ	M	ϰ
	20°		
0.6545	935	0.1022	214
0.4990	805	0.0828	174
0.4055	687	0.0631	143
0.3284	578	0.0513	117
0.2505	466	0.0425	101
0.2036	390	0.0188	49.4
0.1649	325	0.00945	25.8
0.1257	260		
	25°		
0.5621	978	0.0826	191
0.5089	904	0.0714	170
0.4221	785	0.0664	159
0.3281	654	0.0556	143
0.2830	579	0.0414	108
0.2561	524	0.0283	76.9
0.2121	452	0.0216	59.1
0.1643	364	0.0143	41.5
0.1422	319	0.0072	22.1
0.1061	240		
	60°		
1.321	2690	0.2505	92.9
1.090	2540	0.2036	77.7
0.9940	2430	0.1649	64.9
0.8279	2210	0.1257	51.7
0.6545	1900	0.0828	34.6
0.4990	1590	0.0513	24.9
0.4055	1370	0.0425	20.9
0.3284	1140	0.0188	9.85

Water (H_2O) + Sodium thioantimonate (Na_3SbS_4)

Donk, 1908

%	f.t.	%	f.t.
0.5	-0.1	19.3	+15
4.0	0.65	27.1	30
5.7	0.9	32.0	38
7.8	1.26	38.9	49.6
9.2	1.45	45.0	59.6
11.2	1.75	50.7	69.6
11.8	0 9aq.	57.1	79.5

Water (H_2O) + Sodium sulfamate ($NaSO_3NH_2$)

Laning and van der Meulen, 1947

%	f.t.	%	f.t.
44.40	0	62.44	40
52.51	20	63.64	45
57.60	30	64.51	50
60.08	35	65.63	55

Water (H_2O) + Sodium sulfite (Na_2SO_3)

Arii, 1928

t	p 7aq. + 0aq.	p 7aq. + sat.sol.
15	8.76	-
20	12.99	16.24
25	18.59	21.61
30	26.83	28.38
33	32.78	32.99

t	p 0 aq. + sat.sol.	t	p
35	37.01	45	63.74
40	48.88	50	82.39

Kremers, 1856

%	f.t.	
12.39	0	b.t. sat.sol.
22.27	20	105°
33.11	40	

Hartley and Barrett, 1909

%	f.t.	%	f.t.
	ice		
2.10	-0.76	8.62	2.77
4.04	1.37	11.09	3.51
5.92	1.96	15.18	4.50
	7 aq.		
11.57	-1.9	23.03	23.5
12.90	+2.0	25.93	29.0
14.97	5.9	28.39	33.5
16.67	10.6	30.59	37.2
20.19	18.2		
	0 aq.		
22.06	60.4	22.33	59.8
22.27	59.8	22.41	59.8

%	t.spontaneous cryst.	%	t.spontaneous cryst.
ice		salt	
4.80	-2.8	23.03	-10.55
9.02	-4.5	25.75	- 3.5
12.90	-6.15	25.93	- 3.1
16.68	-8.05	26.93	- 0.1
20.20	-9.7	28.49	4.3
		30.59	10.6

Lewis and Rivett, 1924

%	f.t.	%	f.t.
26.8	33	24.7	52.5
26.3	37.5	24.1	57.5
25.7	42.5	23.6	62.5
25.25	47.5	23.2	67.5

Chéneveau, 1907

%	d	n_D	%	d	n_D
		19°			
0	0.9984	1.3331	10.56	1.1004	1.3535
3.74	1.0340	.3404	13.68	.1314	.3601
7.25	.0677	.3474	16.62	.1613	.3658
8.93	.0843	.3508			

Water (H_2O) + Sodium dithionate ($Na_2S_2O_6$)

De Baat, 1926

%	f.t.	%	f.t.
6.05	0	13.39	20
10.63	12	17.32	30

Ishikawa and Okee, 1932

%	f.t.	d	%	f.t.	d
		2aq.			
7.27	0	1.0565	23.18	50	1.1705
8.54	5	-	26.50	60	.1920
9.40	8	-	29.55	70	.2191
10.00	10	1.0773	33.00	80	.2425
13.13	20	.0984	36.02	90	.2711
16.42	30	.1216	39.29	100	.2981
19.79	40	.1460			
		6aq.			
5.90	0	-	7.81	5	1.056
6.04	0.5	-	8.95	7.5	.068
6.21	1	-	9.78	9	-
6.89	3	1.054	10.18	10	-
		8aq.			
5.60	0.73	-	5.86	0.14	-
5.69	0.49	-	5.90	0.00	-
		ice			
1.45	0.66	-	5.13	1.082	-
3.64	0.785	-	5.00	1.136	- E

Water (H_2O) + Sodium carbonate (Na_2CO_3)

Heterogeneous equilibria.

Tammann, 1885

t	p 0%	9.22%	17.93%	21.71%	25.79%
28.68	29.5	28.1	26.9	26.3	25.3
33.63	39.1	37.4	36.1	35.0	33.8
37.81	49.2	47.3	45.5	43.8	42.1
40.19	55.9	54.1	51.3	49.7	48.0
44.73	70.9	68.7	65.3	63.4	60.7
49.01	88.1	84.7	81.2	79.1	75.8
52.35	103.9	100.2	95.9	93.3	89.7
55.63	121.7	118.1	112.2	109.1	105.6
58.70	140.7	135.8	130.0	126.0	122.3
61.06	156.9	151.3	144.7	140.7	135.4
63.53	175.6	170.3	162.2	157.7	153.2
66.07	196.8	190.8	182.6	177.5	171.5
69.30	226.8	219.0	209.6	203.6	197.3
72.33	258.4	249.8	238.8	232.4	225.6
74.75	286.3	276.2	264.2	257.3	249.7
77.84	325.5	314.5	300.7	292.7	284.5
80.36	360.7	348.4	333.6	324.8	315.8
82.88	398.9	385.8	369.3	358.9	349.2
85.96	450.3	435.5	416.8	406.2	395.2
89.21	510.3	491.5	471.5	459.7	445.9
92.63	580.7	560.9	537.0	522.4	509.1
96.47	669.2	645.6	519.3	604.4	589.9
100.08	762.1	733.7	704.6	689.5	671.1

Tammann, 1885

%	p	%	p
	100°		
5.44	744.5	19.09	701.0
10.03	730.5	23.80	681.1
12.68	722.9	29.15	653.6
16.46	711.2		

Nicol, 1884

t	p	t	p
	sat.sol.		
65	154.9	85	364.4
75	239.5	95	536.8

Waldeck, Lynn and Hill, 1932

t	P	t	P
	sat.sol.		
104.8	1.00	231.0	26.6
112.0	1.27	313.0	99.0
121.0	1.69	365.0	199.0
176.5	8.43		

Speranski, 1911

t	p 0%	sat.sol.
		1aq.
24.05	20.30	18.345
25.55	21.94	19.55
27.0	23.74	20.70
28.05	24.61	21.04
28.1	24.56	20.855
28.85	25.51	21.50
29.1	25.68	21.40
30.07	26.71	21.74
36.0	34.93	25.61
39.14	41.54	30.60
41.89	48.39	36.01
43.9	54.11	40.68
46.1	60.80	45.93
47.9	67.02	51.10
49.95	74.35	56.76
		7aq.
25.1	20.22	16.75
26.1	21.48	17.82
27.95	23.47	18.92
28.0	23.96	19.52
29.9	25.99	20.61
29.9	26.07	20.77
29.95	25.99	20.53
33.75	31.19	23.32
35.0	33.09	24.36

Keevil, 1942

mol%	t	P	mol%	t	P
		sat.sol.			
5.3	183.6	9.72	1.4	305.5	89.4
4.2	225.8	24.37	0.5	334.4	131.2
3.4	247.3	35.57	0.1	357.0	179
2.3	278.2	61.0	0.2	368.5	202.7

Taylor, 1955.

molality	p 65°	80°	95°
0	187.6	355.3	634.3
0.1009	186.8	353.5	631.4
.1507	186.4	352.6	629.9
.1999	185.9	351.9	628.7
.4013	184.4	349.2	623.8
.5017	183.6	347.9	-
.5994	182.9	346.6	619.0
.6934	182.2	-	-
.8046	181.5	344.3	614.3
1.001	180.2	342.1	610.2
1.146	179.3	340.5	607.4
1.444	177.3	337.4	601.8
1.519	176.7	336.3	600.4
2.011	173.7	330.8	589.8
2.5120	169.8	324.1	578.2

Lescoeur, 1890

t	p dissociation 10 aq.	p dissociation sat.sol.
10	4.6	-
15	6.9	11.0
20	10.3	16.0
25	14.7	-
30	-	25.2

Gerasimov, 1913

t	p dissociation 1aq.
33.9	29.9
34.5	30.9
37.8	37.9
56.2	99.7
61.0	125.2
75.5	241.0
76.0	246.0
79.7	288.5
100.8	657.3
101.0	661.6

Rode, 1933

Vapour tension of hydrates .

Morey and Chen, 1956

t	P
V + L + C	
374	250
400	1000
500	1000
600	1000

Legrand, 1835

%	b.t.	%	b.t.
0	100.0	26.90	103.0
6.98	100.5	29.08	103.5
12.58	101.0	30.89	104.0
17.21	101.5	32.38	104.5
21.07	102.0	32.66	104.63
24.24	102.5		

Kremers, 1856

b.t. sat.sol. = 106°

Gerlach, 1886

%	b.t.	%	b.t.
0	100.0	23.73	103.0
4.94	100.5	26.58	103.5
9.42	101	29.18	104
13.49	101.5	31.60	104.5
17.21	102	33.86	105
20.63	102.5		

Jones, 1904

M	b.t.	M	b.t.
0.05	100.05	0.5	100.54
0.1	100.13	1.0	100.94
0.2	100.26	2.0	101.73

0.50 N f.t. = -1.882

Freezing curve

Poggiale, 1843

%	f.t.	%	f.t.
6.61	0	23.57	25
14.27	10	26.42	30
20.59	20	32.66	104.6

Loewel, 1851

f.t.	% 10aq.	% 7aq. II	% 7aq. I
0	6.51	24.20	16.94
10	10.76	27.46	20.84
15	13.94	29.36	21.83
20	17.83	31.41	27.83
25	22.18	-	27.58
30	27.14	-	30.29
38	34.06	-	-
104	31.26	-	-

Mulder, 1866

%	f.t.	%	f.t.	
6.63	0	31.88	57	
7.84	4.25	31.60	70.5	
13.19	13.5	30.93 sic	75	
14.23	15.1	31.69	87.4	
18.50	21	31.17	99	
28.47	30.5	31.08	105.0	b.t.
28.72	31	37.11	32.5	10aq. tr.t.
31.51	34			

Ketner, 1902

%	f.t.	
-	31.85	10aq. - 7aq.
31.8	32.1	
32.1	32.5	
32.7	33.3	
33.0	33.9	
33.9	34.5	
-	35.1	7aq. - 1aq.

Wells and Mc Adam, 1907

%	f.t.	%	f.t.
25.48	27.84 (10+1)	33.56	29.86
27.22	29.33	33.66	29.89
28.00	29.85	33.47	31.80
28.63	30.35	33.17	35.17
30.19	31.45	33.18	35.37
30.53	31.66	33.05	35.66
30.65	31.72	33.08	36.45
31.34	32.06	33.04	36.90
30.31	30.35 (7+1)	33.02	37.91
31.11	31.82	32.94	38.92
31.64	32.86	32.93	41.94
32.53	34.37	32.66	40.93
32.88	34.76	32.67	43.94
32.99	35.15	32.42	
33.32	35.17		
33.37	35.62		

tr.t. (10-7) 32.00 (10-1) 32.96 (7-1) 35.37

Hill and Miller, 1927

%	f.t.		%	f.t.	
17.8	10	10aq.	31.6	33	7aq.
22.5	25		33.2	36	1aq.
28.0	30		32.2	50	

Mc Bain, 1927

%	f.t.	
32.8	40	3 aq.
39.8	-	3 aq. - 1 aq.

Seyer and Todd, 1929

%	f.t.	%	f.t.
14.5	16.6	31.3	65.6
16.2	18.4	30.6	79.0
17.8	20.2	30.5	90.0
22.5	25.0	29.8	101.0
28.4	30.0	28.7	122.0
32.8	34.8	28.4	131.0
32.3	46.5	27.6	140.0
32.1	50.4	26.0	161.0
32.0	52.2	24.5	173.0

Waldeck, Lynn and Hill, 1932

%	f.t.	%	f.t.
32.0	50.0	28.1	139.0
31.6	60.0	27.7	145.0
31.0	75.5	27.7	152.0
30.8	83.6	26.8	159.0
30.6	99.5	26.1	170.0
30.8	100.0	25.5	174.0
30.9	100.0	25.5	176.0
30.8	100.5	25.0	180.0
30.8	100.5	23.3	200.0
30.8	103.0	20.7	225.0
30.7	106.0	18.7	239.0
30.8	112.5	13.2	275.0
30.8	113.0	8.4	300.0
30.3	116.0	8.4	301.0
29.8	121.5	4.8	321.0
29.5	125.0	4.4	326.0
28.8	131.5	0.0	348.0

Foote and Vance, 1933

%	f.t.	
6.42	0	10 aq.
22.60	25	
32.83	40	1 aq.
32.16	50	

Deffet, 1836

E	P	E	P
-2.05	1	-7.5	498
-3.50	51	-9.9	756
-5.0	216	-11.5	926

Jones B.M., 1909

spont.cryst.	%	spont.cryst.	%
	ice		
-0.7	0	-6.1	11.91
-1.3	1.48	-6.6	14.16
-3.1	5.04	-7.0	14.21
-5.0	9.30	-7.5	15.48
	10 aq.		
-6.6	16.16	+4.9	28.20
-6.0	18.71	5.0	26.14
-5.3	20.24	5.1	25.50
-3.0	18.20	5.3	26.68
-2.0	19.27	5.9	27.12
-2.0	21.94	6.2	28.03
-0.7	20.06	6.5	28.36
+0.6	22.08	7.2	29.79
+2.0	24.45	7.7	30.99
+2.6	23.41	8.0	32.12
+3.4	24.26	8.4	32.42
+4.0	25.77		
	7 aq.		
1.9	21.33	4.9	26.47
5.0	25.77	6.5	27.12
4.3	26.14		
	1 aq.		
54.5	32.92	69.5	31.75
58.0	32.54	81.0	31.19
60.0	32.30	82.7	31.17
64.0	31.92		

Density

Properties of phases.

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0193	12	1.1254
4	.0403	14	.1476
6	.0622	16	.1702
8	.0823	16.67	.1782
10	.1037		

Schiff, 1858

%	d	%	d
	23°		
0	0.9976	8.04	1.0833
2.01	1.0194	12.06	1.1280
4.02	1.0405	28.08	1.1966
6.03	1.0612		

Gerlach, 1859

%	d	%	d
	15°		
0	0.99913	8	1.08335
1	1.00962	9	.09404
2	.02012	10	.10474
3	.03061	11	.11557
4	.04110	12	.12641
5	.05163	13	.13746
6	.06216	14	.14850
7	.07275		
0.37	1.003	7.40	1.077
0.74	.007	7.77	.081
1.11	.011	8.14	.085
1.48	.015	8.51	.089
1.85	.019	8.88	.093
2.22	.022	9.25	.098
2.59	.026	9.62	.102
2.96	.030	9.99	.105
3.33	.034	10.36	.109
3.70	.038	11.10	.113
4.07	.042	11.47	.118
4.44	.046	11.84	.122
4.81	.049	12.21	.129
5.18	.053	12.58	.134
5.55	.057	12.95	.138
5.92	.061	13.33	.142
6.29	.065	13.69	.146
6.66	.069	14.06	.149
7.03	.075		

Favre and Valson, 1874

m	d	m	d
	23.3°		
0	0.9975	1.5	1.145
0.5	1.052	2.0	1.187
1.0	1.100	2.5	1.226

Kohlrausch, 1879

%	d
18°	
5	1.0511
10	1.1044
15	1.1590

Lunge, 1882

%	d	%	d
	30°		
0	0.996	21.33	1.230
13.62	1.140	22.21	1.240
14.47	1.150	23.08	1.250
15.32	1.160	23.93	1.260
16.18	1.170	24.78	1.270
17.04	1.180	25.62	1.280
17.90	1.190	26.46	1.290
18.76	1.200	27.30	1.300
19.61	1.210	28.13	1.310
20.47	1.220		

Volkmann, 1882

%	d
14° - 15°	
12.58	1.1329
5.83	1.0605
2.76	1.0283

Moore, 1895-96

M	d	M	d
	18°		
0.00	0.9987	1.0	1.0980
0.25	1.0250	2.0	1.1880
0.5	1.0517		

Gilbault, 1897

%	d	%	d
	20°		
0	0.9982	9	1.0931
5	1.0506	14.028	1.1478

Pann, 1901

%	d 10°	d 18°	d 30°
0	0.999727	0.998622	0.995673
5.00	1.05306	1.05076	-
6.60	1.07096	1.06830	-
10.50	1.10779	1.09944	1.10551
24.40	-	1.26570	1.19371
63.1	-	-	1.28109

Wegscheider and Walter, 1905 and 1906

%	d	%	d
		60°	
0	0.9832	22.25	1.2191
14.06	1.1277	25.20	1.2546
18.23	1.1746	28.74	1.2971
		80°	
0	0.9718	28.59	1.2807
18.26	1.1607		

Lunge and Berl, 1907

%	d 15°	d 25°	%	d 15°	d 25°
2.00	1.022	1.018	10.85	1.116	1.111
4.16	1.045	1.041	13.25	1.142	1.136
6.36	1.067	1.063	14.09	1.152	1.146
8.57	1.091	1.087			

Wasastjerna, 1920

M	d	M	d
20°		25°	
0.0000	0.99823	0.0000	0.99707
.1076	1.01017	.1075	1.00886
.1982	.01993	.1979	.01856
.2914	.02971	.2910	.02819
.3962	.04061	.3956	.03897
.4880	.04997	.4872	.04828
.5940	.06070	.5930	.05895
.6933	.07057	.6921	.06871
.7914	.08024	.7900	.07835
.8519	.08610	.8504	.08414
.9883	.09924	.9864	.09718
1.0951	.10942	1.0930	.10729
.1849	.11792	.1826	.11578
.2544	.12438	.2519	.12216
.3809	.13602	.3781	.13371
.4816	.14518	.4786	.14284
.5765	.15376	.5733	.15141

Rakshit, 1925

%	d	%	d
	20°		
1	1.00791	10	1.09016
5	1.04604	30	1.25032

Faust, 1927

N	%	d
	20°	
1.03	5.22	1.055
2.22	10.60	1.111
3.935	17.54	1.1891

Hrynakowski, 1927

Sat. sol. 50.00° d = 1.3414

Flöttmann, 1928

%	t	d
14.226	15	1.1515
18.02	20	1.1941
22.55	25	1.2416

Okazaki, 1933

%	d	%	d
	28°		
5.59	1.0546	15.91	1.1669
9.12	.0908	18.79	.1996
14.84	.1545		

Pesce, 1934

N	d	N	d
	25°		
0.948	1.04671	3.560	1.16887
1.323	.06535	3.639	.17221
1.834	.09000	4.817	.22260
2.626	.12690		

Hitchcock and Mc Ilhenny, 1935

N	d	N	d	N	d
20°		30°		40°	
0.9644	1.0481	0.9990	1.0472	0.9583	1.0407
1.945	.0954	2.950	.1372	1.930	.0872
2.850	.1375	3.937	.1826	2.828	.1281
3.845	.1848	5.230	.2367	3.814	.1717
4.824	.2240	5.974	.2673	4.786	.2137
5.684	.2591	6.989	.3062	5.639	.2438
6.620	.2946	7.772	.3362	6.570	.2838
7.618	.3308			7.533	.3199

Roberts and Mangold Jr., 1939

t	d		
	1 N	2 N	3 N
22.5	1.047	1.095	1.141
40.0	1.043	1.090	1.134
50.0	1.038	1.084	1.128
60.0	1.032	1.077	1.122
70.0	1.026	1.072	1.116
75.0	1.023	1.068	1.112
80.0	1.020	1.065	1.109
85.0	1.017	1.062	1,105
	4 N	5 N	6 N
22.5	1.185	1.225	1.262
40.0	1.177	1.217	1.254
50.0	1.170	1.211	1.249
60.0	1.164	1.205	1.243
70.0	1.158	1.198	1.235
75.0	1.154	1.194	1.231
80.0	1.152	1.190	1.228
85.0	1.148	1.187	1.224

Grassi, 1851

17.185 % 22.2° π = 29.7

Gilbault, 1897

%	π	%	π
	20°		
0	44.37	9	33.28
5	37.77	14.028	28.42

Viscosity and surface tension

Slotte, 1883

%	η 20°	η 30°	η 40°
5.96	1371	1088	887
11.5	1963	1529	1229
15.85	2773	2116	1658

Moore, 1895-96

18°

M	η (H_2O=1)	M	η (H_2O=1)
0.00	1.000	1.0	1.667
.25	1.120	2.0	3.128
.50	1.274		

Faust, 1927

20°

N	%	η
1.03	5.22	1140
2.22	10.60	1440
3.935	17.54	2210

Hrynakowski, 1927

sat.sol. 50.00° η= 5517

Hitchcock and Mc Ilhenny, 1935

N (20°)	η (20°)	N (30°)	η (30°)	N (40°)	η (40°)
0.9644	1272.5	0.9990	1036.2	0.9583	838.7
1.945	1669.3	2.950	1787.4	1.930	1068.4
2.850	2200.2	3.937	2405.6	2.828	1364.2
3.845	2070.5	5.230	3649.5	3.814	1807.6
4.824	4351.5	5.974	4771.0	4.786	2424.8
5.684	6015.8	6.989	6858.3	5.639	3158.3
6.620	8695.0	7.772	9393.6	6.570	4184.0
7.618	13228.0			7.533	5732.4

Roberts and Mangold Jr, 1939

t	η 0%	η 1N	η 2N	η 3N	η 4N	η 5N	η 6N
22.5	949	1201	1564	2116	2863	3980	5630
50.0	549	756	976	1283	1678	2246	2997
75.0	380	468	584	746	950	1198	1521
98.0	289	347	428	530	647	794	960

Volkmann, 1882

14° - 15°

%	σ
12.58	76.3
5.83	74.9
2.76	74.3

Pann, 1901

%	σ 10°	σ 30°
0	74.01	71.01
5.0	75.34	-
6.60	75.62	-
10.50	76.87	73.76
24.40	-	76.43
63.1	-	80.44

Forch, 1905

M	t	σ
2.728	15.8	80.23
2.198	15.6	79.55
1.754	15.5	79.00
1.462	15.2	78.64
1.091	15.3	78.13

Faust, 1927

20°

N	%	σ
1.03	5.22	73.7
2.22	10.60	74.9
3.935	17.54	77.3

Hrynakowski, 1927

sat.sol. 50.00° σ= 82.92

Börner, 1869

t	nH_α	t	nH_β	t	nH_γ
		10%			
45.1	1.346749	44.6	1.353228	43.7	1.356993
40.85	.347485	40.3	.353942	40.1	.357492
35.3	.348434	34.95	.354866	34.65	.358435
30.5	.349187	30.6	.355549	30.35	.359147
		20%			
45.1	1.362454	44.1	1.369364	42.2	1.373320
40.4	.363235	40.2	.370107	38.5	.374060
35.6	.364122	35.3	.370884	34.9	.374624
29.55	.365204	29.4	.372014	29.3	.375716
		30%			
44.9	1.375012	45.3	1.382007	44.9	1.385965
39.1	.376174	39.1	.383129	41.3	.386524
33.9	.377054	33.3	.384320	32.9	.388341
28.3	.370144	28.3	.385230	27.5	.389214

Jones, 1904 and Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.05	1.32601	0.50	1.33584
0.10	.32735	1.00	.34536
0.20	.32957	2.00	.36209

Wasastjerna, 1920

M	n 6563 Å	n 5893 Å	n 4861 Å
		20°	
0.0000	1.33151	1.33330	1.33747
.1076	.33405	.33587	.34008
.1982	.33610	.33795	.34220
.2914	.33813	.34000	.34428
.3962	.34037	.34226	.34656
.4880	.34228	.34420	.34853
.5940	.34445	.34641	.35077
.6933	.34644	.34841	.35280
.7914	.34840	.35037	.35478
.8519	.34954	.35153	.35598
.9883	.35214	.35413	.35863
1.0951	.35412	.35613	.36068
.1849	.35570	.35773	.36234
.2544	.35701	.35903	.36366
.3809	.35913	.36113	.36580
.4816	.36095	.36298	.36766
.5765	.36254	.36459	.36931
		25°	
0.0000	1.33108	1.33290	1.33700
.1075	.33356	.33529	.33955
.1979	.33556	.33732	.34162
.2910	.33756	.33934	.34367
.3956	.33973	.34164	.34591
.4872	.34164	.34355	.34790
.5930	.34386	.34575	.35013
.6921	.34572	.34769	.35213
.7900	.34763	.34961	.35405
.8504	.34876	.35077	.35529
.9864	.35140	.35336	.35790
1.0930	.35328	.35532	.35993
.1826	.35492	.35699	.36151
.2519	.35615	.35819	.36186
.3781	.35838	.36042	.36503
.4786	.36014	.36211	.36687
.5733	.36179	.36710	.36850

Flöttmann, 1928

%	t	n_D
14.226	15	1.36422
18.02	20	.37126
22.55	25	.37973

Kohlrausch, 1879

%	κ	$\tau.10^4$
	18°	
5	448	253
10	700	272
15	831	295

Jones, 1904

M	λ	M	λ
	0°		
0.05	85.80	0.50	53.10
0.10	76.80	1.00	39.14
0.20	66.90	2.00	24.53

Heim, 1886

t	R (arbitrary units)	t	R (arbitrary units)
	21.29%		
50	88.0	25	160.2
48	91.3	20	186.1
44.5	98.2	15	218.6
40	109.8	10	260.2
35	123.0	5	316.4
30	139.6		

Roberts and Mangold Jr, 1939

t	κ 1 N	κ 2 N	κ 3 N
25.0	530.5	808.8	965.6
35.0	649.0	9997.0	1211.0
	4 N	5 N	6 N
25.0	1027.0	1016.0	952.6
35.0	1311.0	1298.0	1250.0

molarity	e.f. $.10^5$	molarity	e.f. $.10^5$
	25°		
1.000	100812	4.000	103489
2.000	102099	5.000	103804
3.000	102931	6.000	103903

e.f. = electromotive force

Okazaki, 1933

%	Verdet's constant (3441 Å)
28°	
5.59	0.04508
9.12	.04576
14.84	.04698
15.91	.04714
18.79	.04775

Heat constants.

Marignac, 1876

mol%	U(21-52°)	mol%	U(21-52°)
4	0.8649	1	0.9435
2	0.9072	0.5	0.9695

Swallow and Alty, 1931

%	U 17.6°	U 30.0°	U 76.6°	U 98.0°
0.000	0.9992	0.9986	1.0098	1.0084
1.498	.9807	-	-	-
2.000	-	0.9786	-	-
2.901	0.9597	-	-	-
4.000	-	0.9594	-	-
5.000	0.9428	-	0.9761	-
6.000	-	0.9392	-	-
8.000	0.9183	-	-	-
10.000	0.9086	-	0.9452	-
13.790	0.8924	-	-	-
13.840	-	0.8881	-	-
20.000	-	0.8631	0.8936	-
25.000	-	-	0.8615	0.8911

% initial	% final	Q dil.
	30°	
28.24	14.04	543.6
14.04	7.43	518.1
7.43	3.87	399.7
3.87	1.97	342.8
28.24	14.13	525.0
14.13	7.52	582.9
7.52	3.86	401.5
3.86	1.95	346.2
28.40	14.96	501.1
14.96	7.53	583.7
7.53	3.82	426.4
10.00	5.05	426.0

Water (H_2O) + Sodium acid carbonate ($NaHCO_3$)

Poggiale, 1843

%	f.t.	%	f.t.
7.34	0	10.52	40
8.16	10	11.28	50
8.96	20	12.03	60
9.77	30	12.77	70

Dibbits, 1874

%	f.t.	%	f.t.	%	f.t.
6.45	0	8.87	21	11.42	41
6.54	1	9.00	22	11.51	42
6.63	2	9.13	23	11.66	43
6.71	3	9.25	24	11.81	44
6.84	4	9.38	25	11.93	45
6.93	5	9.50	26	12.09	46
7.06	6	9.62	27	12.20	47
7.15	7	9.74	28	12.36	48
7.28	8	9.87	29	12.51	49
7.41	9	9.99	30	12.62	50
7.44	10	10.11	31	12.77	51
7.62	11	10.23	32	12.92	52
7.75	12	10.35	33	13.04	53
7.88	13	10.50	34	13.19	54
8.00	14	10.66	35	13.34	55
8.14	15	10.72	36	13.49	56
8.26	16	10.87	37	13.64	57
8.38	17	11.02	38	13.79	58
8.50	18	11.11	39	13.94	59
8.62	19	11.27	40	14.09	60
8.74	20				

Fedotiev, 1904

%	f.t.	%	f.t.
40.8	0	52.4	30
46.8	15	58.1	45

Nishizawa, 1920

m	f.t.
18.53	15.0
23.61	30.0
26.92	40.0

Neumann and Domke, 1928

c	f.t.
9.08	20
10.42	30
11.96	40

Waldeck, Lynn and Hill, 1934

%	f.t.	%	f.t.
19.1	100	37.5	190
27.2	150	43.0	200
32.0	170		

Hrynakowski, 1927

Sat sol. 50.65° $d = 1.1294$

$\eta = 1101$

$\sigma = 68.94$

Water (H_2O) + Sodium sesquicarbonate ($Na_3HC_2O_6$)

Poggiale, 1843

%	f.t.	%	f.t.
11.21	0	22.89	60
13.42	10	24.51	70
15.47	20	26.36	80
17.45	30	27.87	90
19.32	40	29.38	100
21.12	50		

Water (H_2O) + Sodium metasilicate (Na_2SiO_3)

Baker and Jue, 1938

%	f.t.	
2.21	-1.00	
5.6	-2.7	E
6.29	-1.00	9 aq.
6.63	0.60	
15.80	20.00	
18.20	25.00	
21.70	29.90	
25.66	35.00	
30.18	39.90	
36.18	45.00	
42.94	47.85	9 aq. m.t.
42.2	47.6	9 aq. - 8aq.
41.3	46.7	8 aq.
45.85	48.35	8 aq. m.t.
39.80	46.8	9 aq. - 6 aq.
39.70	47.00	
42.00	52.00	
44.30	56.00	
48.31	60.00	
53.03	62.85	6 aq. m.t.
46.50	55.00	
48.25	61.00	
48.0	59.8	6 aq. - 5 aq.
50.05	65.00	
52.40	70.00	
54.60	71.30	
57.53	72.20	5 aq. m.t.
47.97	90.00	5 aq. - 0aq.
49.17	85.00	
51.63	80.00	
54.75	74.00	
55.75	74.00	
56.6	72.0	

Traube, 1895

%	d	%	d
	20°		
0	0.9982	19.025	1.21598
4.872	1.04992	26.724	1.31744
12.673	1.13822		

Rubien, 1911

N	d	N	d
	18°		
0.1001	1.00531	0.9908	1.06151
0.2024	1.01198	1.967	1.11960
0.5037	1.03130	3.920	1.22922

Heydweiller, 1912

N	d	N	d
	18°		
0.0991	1.00531	1.947	1.11960
0.2004	.01198	3.88	1.22922
0.499	.03130	5.88	1.33389
0.981	.06151		

Main, 1926

N	d	N	d
	25°		
2.435	1.1458	0.841	1.0529
1.134	1.0705	0.546	1.0352

N	η	N	η
	25°		
2.435	2158	0.841	1256
1.134	1381	0.546	1159

Rubien, 1911

N	n_D	N	n_D
	18°		
0.1001	1.33474	0.9908	1.34645
0.2024	1.33618	1.967	1.35819
0.5037	1.34024	3.920	1.37923

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34657
0.1	.33474	2.0	1.35859
0.2	.33615	4.0	1.38011
0.5	.34019		

Kohlrausch, 1892

N	ϰ 18°	$\tau.10^4$(18 - 26°)
0.0001	0.133	272
0.001	1.42	232
0.01	13.6	214
0.1	115	220
0.5	432	236
1	702	244
2	1020	263
3	1113	288
4	1074	335
6	777	-

Heydweiller, 1912

N	ϰ	N	ϰ
	18°		
0.0991	114.2	1.947	1016
0.2004	210.2	3.88	1095
0.499	437.1	5.88	805
0.981	706.1		

XlVI. WATER + SODIUM POLYOXY AND ORGANIC SALTS

Water (H_2O) + Sodium sulfate (Na_2SO_4)

Heterogeneous equilibria .

Wüllner, 1885

t	p 0%	4.76%	9.09%	13.04%	16.67%	20%
26.3	25.43	25.03	24.73	24.23	23.83	-
27.9	27.93	-	27.08	26.58	-	-
28.7	29.27	28.68	28.19	27.69	27.29	26.90
32.2	35.76	35.07	34.57	33.98	33.49	32.99
35.7	43.43	42.69	41.99	41.30	40.71	40.21
39.5	53.46	52.68	51.89	51.09	50.19	49.50
40.2	54.50	-	-	52.02	51.53	50.84
42.7	63.35	62.47	61.67	60.88	60.09	59.25
44.7	70.31	69.23	68.23	67.33	-	65.49
45.4	72.58	-	70.50	69.31	68.52	67.73
47.2	79.90	78.80	77.80	76.90	75.90	75.00
47.6	81.53	80.35	79.35	78.36	77.27	76.18
49.6	90.18	-	-	-	85.48	84.54
50.6	93.88	-	-	90.22	89.13	87.84
52.5	104.01	102.63	101.14	99.95	98.67	97.28
55.1	118.04	116.46	114.84	-	111.53	109.89
57.1	129.88	128.30	126.81	125.22	123.64	122.36
58.4	138.14	136.66	135.08	133.39	131.91	130.43
60.6	152.61	150.83	148.99	146.98	145.39	143.81
61.7	160.96	158.99	156.76	154.57	152.51	-
63.9	177.92	-	173.58	171.20	169.62	167.54
65.8	193.74	-	-	185.84	183.66	181.74
67.3	205.13	202.66	199.98	197.42	195.04	192.68
69.3	231.10	-	-	222.69	220.04	217.48
72.5	259.57	256.12	253.15	250.09	246.74	243.77
73.9	275.47	-	269.26	266.00	263.04	260.08
76.2	303.37	300.02	296.07	292.69	289.45	286.21
78.7	336.38	-	-	324.36	320.90	317.45
79.5	347.53	343.43	339.25	336.31	331.76	327.42
81.1	370.79	366.49	362.03	357.59	353.15	-
84.0	416.29	-	406.84	401.81	398.27	393.44
86.5	459.21	-	447.99	442.18	436.81	430.81
86.8	464.64	459.14	453.44	447.92	442.50	-
87.1	470.04	-	459.12	453.70	448.68	443.26
89.9	423.45	517.35	511.57	504.85	499.05	491.95
91.0	545.78	-	533.63	527.68	521.68	515.48
93.2	592.82	585.98	578.47	571.97	564.87	558.31
95.8	657.73	649.59	642.50	634.73	628.83	619.69
96.9	679.55	671.75	664.35	-	-	-
100.1	761.40	-	-	732.70	723.80	714.40
100.2	764.80	755.27	746.43	736.41	728.25	-
100.6	776.71	-	758.05	749.21	740.37	731.04

Tammann, 1885

%	p	%	p
	100°		
4.81	751.0	24.74	703.5
11.16	737.9	29.80	686.4
17.98	721.8		

t	p 0%	11.74%	16.82%	27.19%	32.81%
41.59	60.2	58.5	57.0	54.8	53.7
44.56	70.3	68.3	66.4	63.9	62.5
48.48	85.8	83.2	81.3	77.9	76.0
50.49	94.8	91.7	90.0	87.1	83.8
56.41	126.5	122.9	120.4	115.4	113.0
58.52	139.5	135.2	132.4	127.0	114.4
60.02	149.6	144.1	142.3	136.2	132.9
64.87	186.5	180.3	177.0	169.3	166.0
68.47	218.8	211.7	207.8	198.4	195.0
70.58	239.7	232.7	228.1	218.1	213.7
75.33	293.4	284.0	278.7	266.4	261.6
78.08	328.8	319.1	313.5	299.8	294.0
80.91	368.8	352.4	351.2	335.3	329.1
83.82	414.2	400.6	393.3	376.5	369.4
86.88	457.9	444.0	436.1	416.9	409.4
88.92	504.9	489.1	480.8	460.0	451.3
91.91	548.5	517.3	522.2	500.2	490.2
93.61	602.2	589.3	574.8	550.3	539.4
97.38	691.4	669.7	657.8	630.3	619.3
100.35	769.5	744.8	732.2	702.6	670.7

Perreu, 1930

%	p	%	p
	30°		
50.0	29.61	60.00	28.76
55.56	29.19	62.96	28.43

Foote, Saxton and Dixon, 1932

$$\lg p = -\frac{2696.6}{T} + 10.3630$$

Pearce and Eckstrom, 1937

M	p	M	p
	25°		
0.0	23.752	0.8	23.072
0.1	23.652	1.0	22.927
0.2	23.567	1.5	22.565
0.4	23.393	1.9641	22.239
0.6	23.227		

Lescoeur, 1890

t	p (sat.sol.) (10+1)	(7+1)	p 10aq.-7aq.	7aq.-0aq.
0	-	-	3.8	-
5	-	5.5	5.1	-
10	8.7	8.0	7.0	-
15	12.0	10.8	9.7	7.0
20	15.7	15.0	13.9	10.5
25	20.0	19.1	19.0	16.1
29	-	-	24.0	-
30	24.8	23.8	-	22.3

Smits and Wuite, 1909 -1910

t	p	t	p
	C + L + V		
19.4	13.9	56.8	114.0
19.8	14.2	59.6	130.0
20.6	14.9	62.8	151.0
21.6	15.9	65.6	171.0
22.6	17.0	69.4	200.0
23.5	18.0	69.8	208.0
24.6	19.2	74.8	258
25.8	20.6	80.2	321
28.1	23.9	86.0	403
30.5	27.6	90.6	482
32.4	30.8	95.2	574
32.8	31.4	100.0	681
33.8	33.5	105.0	812
35.0	36.0	111.0	1009
37.0	40.1	116.1	1196
39.0	44.6	119.8	1295
40.4	48.2	124.8	1569
41.7	52.3	125.8	1617
44.7	61.9	129.8	1833
48.6	75.5	135.6	2172
50.6	83.6	140.8	2521
53.8	99.0		
	0 aq. + L + V (P kg)		
141.9	3.5	181.4	9.6
142.2	3.6	184.1	10.3
144.7	3.7	186.1	10.7
145.0	3.9	187.6	11.0
147.2	4.1	192.6	12.3
148.2	4.2	201.3	14.2
149.0	4.2	206.4	16.5
151.6	4.5	210.2	18.2
151.7	4.6	217.2	20.4
152.0	4.7	220.0	21.2
153.2	4.9	222.8	22.2
154.2	5.0	226.4	23.9
157.0	5.3	230.8	25.6
160.3	5.7	233.6	27.1
162.5	6.1	234.0	27.1
163.4	6.2	237.2	28.9
166.9	6.8	240.3	31.0
170.4	7.3	244.7	33.7
172.4	7.7	247.0	35.2
174.4	8.1	250.4	38.5
177.3	8.6	255.2	41.4
179.4	9.2	263	45
	7 aq. + L + V		
19.1	14.5	21.9	16.9
20.0	15.3	23.0	17.9
21.1	16.1		
	10 aq. + L + V		
21	16.4	26.1	21.8
22.4	17.8	27.0	23.0
22.6	18.1	28.2	24.7
23.0	18.5	29.4	26.3
23.2	18.7	29.9	27.0
23.6	19.2	31:0	28.6
25.0	20.6	32.4	30.8
	7 aq. + 0 aq. + V		
20.0	13.2	24.1	18.3
21.0	14.3	24.4	18.9
22.0	15.4	24.6	19.2
22.9	16.6	24.9	19.6
23.6	17.6	25.3	20.0
23.9	18.0		
	10 aq. + 0 aq. + V		
20.2	12.5	27.0	21.0
21.6	14.1	28.1	22.7
23.0	15.7	29.5	25.0
24.1	17.0	31.0	27.9
25.0	18.1	32.4	30.8
25.9	19.4		
	10 aq. + 0 aq. + L (P kg)		
32.645	100	32.557	1000
32.679	200	31.41	2000
32.722	460	29.17	3000

Applebey and Hughes, 1915

t	p	t	p
	sat.sol.		
104.245	799.1	103.34	767.1
103.78	789.35	102.28	743.0
103.21	770.0	101.33	715.8

Matsui and Nonatsui,etc., 1925

t	p(10aq.-0aq.)	t	p(10aq.-0aq.)
20	13.44	33	34.98
25	19.86	35	39.41
27	23.57	38	45.69
30	28.95	40	49.87
32	32.27	45	64.55
32.5	34.11		

Pohle, 1927

t	p	t	p
	sat.sol.		
30.0	31.0	65.0	168.0
35.0	37.0	70.0	213.0
40.0	47.5	75.0	260.0
45.0	61.5	80.0	322.0
50.0	82.0	85.0	392.0
55.0	106.0	90.0	473.0
60.0	134.0	95.0	576.0

Arii, 1928

t	p	t	p
	sat.sol.		
35	37.01	45	63.74
40	48.88	50	82.39

Leopold and Johnston, 1927

m	t	p	m	t	p
1.427	20.83	16.01	3.473	35.02	36.74
1.953	25.02	20.44	3.371	42.88	57.03
2.552	28.46	25.17	3.283	50.55	84.39
3.473	32.38	31.47			

Keevil, 1942

mol%	t	p	mol%	t	p
		sat. sol.			
5.1	158.8	5.4	2.9	292.5	73.0
5.4	209.9	17.0	1.1	323	116.4
5.5	227.0	25.1	0.2	355.5	172.3
5.0	258.5	41.0	0.1	366.8	201.0

Morey and Chen, 1956

t	P
V+L+C	
374	250
400	839
500	1000

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
0.1	0.793	1.2	0.631
0.2	0.753	1.4	0.625
0.3	0.725	1.6	0.621
0.4	0.705	1.8	0.620
0.5	0.690	2.0	0.621
0.6	0.678	2.5	0.635
0.7	0.667	3.0	0.661
0.8	0.658	3.5	0.696
0.9	0.650	4.0	0.740
1.0	0.642		

Rode, 1933

Vapour pressure of hydrated solutions.

Gerlach, 1886

%	b.t.	%	b.t.
0	100	24.81	102
8.68	100.5	28.06	102.5
15.25	101	30.79	103
20.63	101.5	31.33	103.2

Jones and Getman, 1904

N	p	b.t.	N	p	b.t.
0.05	763.0	100.05	1.0	763.5	101.06
0.1	763.1	100.11	1.5	765.0	101.74
0.2	763.3	100.23	2.0	764.8	102.08
0.5	763.5	100.53			

1 N f.t. = -1.839

Gerlach, 1926

%	b.t.	%	b.t.
1.9	100.2	6.2	100.6
3.2	100.3	10.8	100.8
3.8	100.4	19.3	101.4

Berkeley and Applebey, 1911

sat.sol. b.t.= 102.885°

Freezing curve .

Loewel, 1851

%	f.t.	%	f.t.	%	f.t.
10aq.		7aq.		0aq.	
4.78	0.00	16.40	0.00	29.90	103.17
8.51	10.00	23.37	10.00	30.06	84.42
10.07	13.00	25.52	13.00	30.72	70.61
11.66	15.00	27.77	15.00	31.23	59.79
12.51	16.00	27.92	16.00	31.89	50.40
13.49	17.00	28.57	17.00	31.34	45.04
14.38	18.00	29.40	18.00	32.79	40.15
15.32	19.00	30.24	19.00	33.01	36.00
15.25	20	34.62	25	33.12	34.00
21.88	25	35.47	26	33.20	33.00
23.08	26			33.50	30.00
28.96	30			33.91	26.00
33.66	33			34.00	25.00
35.48	34			34.54	20.00
				34.75	18.00

Rüdorff, 1862

%	f.t.	%	f.t.
0.99	-0.3	2.91	-0.9
1.96	-0.6	3.85	-1.15

Mulder, 1866

%	f.t.	%	f.t.
4.58	0	33.29	35
6.10	4.25	33.02	39
7.58	9.5	30.12	79
9.34	12.4	29.68	103.5
23.20	26.75		

tr.t.= 33°

de Coppet, 1872

%	f.t.	%	f.t.
1.96	-0.6	9.09	-2.75
1.99	-0.6	9.21	-2.7
3.85	-1.2	10.87	-3.05
3.90	-1.2	13.04	-3.65
4.76	-1.4	13.19	-3.65
4.83	-1.45	16.67	-4.5
5.74	-1.7		

Tilden and Shenstone, 1884

%	f.t.	%	f.t.
29.78	100	30.02	160
29.56	120	30.67	180
29.58	140	31.69	230

Nicol, 1885

%	f.t.	%	f.t.
33.81	33	35.63	25
34.19	30	35.74	20

Etard, 1894

%	f.t.	%	f.t.
4.1	0	30.0	83
6.2	7	29.7	99
9.9	13	29.4	134
19.3	24	29.8	150
25.2	28	29.9	190
29.5	30	30.0	240
32.8	49	24.5	279
31.3	62	17.8	320

Löwenherz, 1895

%	f.t.	%	f.t.
21.62	24.83	33.09	34.85
26.48	28.32	32.58	39.92
32.34	31.90		

Dawson and Williams, 1900

%	f.t.		%	f.t.
32.35	31.94	10 aq.	33.13	33.23
34.26	32.93		32.84	36.22
33.23	32.06	0aq.	32.20	38.31

tr.t.= 32.38

Hartley, Jones and Hutchinson, 1900

%	f.t.	%	f.t.		%	f.t.	
0	-0.5	20.25	-5.9	7aq.	30.13	97.8	0aq
0.99	-0.9	20.62	-5.3		30.19	97.6	
2.52	-1.3	22.16	-3.6		30.59	94.0	
5.61	-2.3	23.23	-1.4		31.41	86.4	
9.36	-3.5	24.35	0.2		31.66	81.7	
10.63	-3.8	24.98	1.0		32.01	80.0	
12.30	-4.5	25.45	1.9		32.72	69.8	
15.55	-5.0	26.42	3.0		33.18	65.8	
17.99	-5.7	26.63	3.0		33.41	64.3	
18.87	-5.7	27.43	4.8		33.73	60.7	
20.25	-6.2	27.59	5.9		33.86	56.7	
		28.71	7.3				

Seidell, 1902

%	f.t.	%	f.t.
8.38	10	28.43	30
17.81	21.5	32.65	33
22.32	25	32.40	35
23.74	27		

Schreinemakers, 1910-11

mol%	f.t.	
sat.sol.		
5.894	35	
3.43	25	10 aq.
1.64	15	

Massink, 1913

%	f.t.
8.26	10
16.25	20
21.88	25

Wuite , 1914

f.t.	mol%	f.t.	mol%
62	5.39	192	5.27
70	5.27	208	5.39
72	5.255	241	5.39
80	5.18	250	5.01
120	5.04	279	4.12
190	5.255	319	2.56

Nishizawa, 1920

%	f.t.
11.52	15
29.08	30
32.43	40

Maass and Hatcher, 1922

%	f.t.	%	f.t.
0	-1.72	13.51	- 6.77
1.19	-2.17	16.77	- 9.02
2.36	-2.87	19.47	-10.57
4.47	-3.12	22.64	+19.6
6.92	-4.27	23.59	29.1
10.23	-5.52	25.62	39.1

Pascal and Ero, 1923

25.90 % f.t. = 14° 7 aq.

Foote , 1925

%	f.t.
33.10	35
21.75	25

Küpper, 1926

%		f.t.	%	f.t.
4.22	10aq.	0	31.85	50
6.23		5	31.55	55
8.24		10	31.43	57
11.96		15.3	31.31	59
14.25		17.9	31.25	60
16.10		20	31.19	61
16.80		20.6	31.03	64
21.90		25	30.70	70
24.78		27	30.66	71
29.10		30	30.30	80
30.74		31	30.19	83
33.04		35.5	29.95	90
32.50	0 aq.	40	29.90	100
31.93		49		

Richards and Ungve, 1928

%	f.t.		%	f.t.
21.55	24.845	10aq.	13.65	17.472
18.43	22.265		11.44	14.731
18.87	19.860			

Matsui, Oguri and al., 1929

%	f.t.	%	f.t.
23.08	26.00	32.00	33.00
25.64	28.00	32.89	35.00
28.71	30.00	32.43	37.00
30.73	31.00	32.44	39.00

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
30.1	249	22.9	290
29.4	256	21.2	295
29.3	258	19.7	300
27.6	271	16.9	306
26.3	277	16.3	308
25.4	281	7.8	325

Eddy and Menzies, 1940

mol%	f.t.	
3.36	24.8	10 aq.
2.58	2.9	7 aq. (metast.)
3.36	8.1	
4.22	13.4	
4.98	17.8	
5.59	20.8	
5.67	21.4	
5.87	24.1	
5.94	21.0	0 aq. (metast.)
5.59	47.5	(stable)
5.07	82.2	
5.00	114.0	
5.02	134.5	
5.07	165.8	

Benrath, 1941

%	f.t.	%	f.t.
28.2	262	15	299
25	275	10	308
20	289	5	319

Booth and Bidwell, 1950

%	f.t.	%	f.t.
5.12	330 ±1	0.87	360 ±3
1.43	354 ±2	0.34	382 ±2

Möller, 1868

P	%		
f.t.	0	15	15.4
1	4.40	11.32	11.44
20	4.53	10.60	10.74
30	-	10.05	-
40	-	10.33	-

Styrikovitch and Khaibullin, 1956 (fig.)

%	f.t.	%	f.t.
5.7	0	33	200
38	20	29	300
33	100	1	370

mg/kg(V)	t	mg/kg(V)	t
160kg/cm_2		200 kg/cm_2	
0.003	350	0.4	370
0.002	380 sat.	0.09	400
		0.08	460 sat.
180 kg/cm_2			
0.1	360		
0.06	380		
0.05	420 sat.		

crit.p.: 5 mg/kg 220 kg/cm_2 480°

Schroeder, Berk and Partridge, 1937

%	P	f.t.	%	P	f.t.
30.7	13.6	197	31.27	24.5	228
30.7	13.6-14.0	196	27.5	55.4	272
30.7	14.0-14.2	201	27.5	55.8	272
30.7	13.6	200	27.5	55.8	273
30.7	13.6	199	15.15	95.2	309
31.25	21.7-22.5	224	15.15	95.2	310
31.25	22.5	225	2.50	156.5	347
31.25	23.8	225	2.50	158.2	347
31.25	24.5	226			

Keevil, 1942

mol%	f.t.	mol%	f.t.
5.1	158.8	2.9	292.5
5.4	209.9	1.1	323
5.5	227.0	0.2	355.5
5.0	258.5	0.1	366.8

Properties of phases.

Density

Bischof, 1850

%	d	%	d
18.75%			
2	1.0167	10	1.0915
4	.0351	12	.1109
6	.0537	14	.1305
8	.0725		

Kremers, 1855

%	d	%	d
19.5°			
0	0.9983	7.995	1.0714
2.894	1.0245	10.539	.0958
5.589	1.0490	12.473	.1143

t	d 1	d 2	% from table above
19.5	1.0610	1.1146	
40	.0531	.1055	
60	.0408	.0949	
80	.0314	.0827	
100	.0181	.0694	

Buchholz, 1856

%	d
18.75°	
12	1.1107
14	.1301
15.50	.1454

Schiff, 1858 and 1859

%	d	%	d
19°			
0	0.9984	5.880	1.0516
1.470	1.0115	8.820	.0789
2.940	.0247	13.230	.1205
4.410	.0382		
0.44	1.0024	3.97	1.0342
0.88	.0063	4.41	.0382
1.32	.0102	4.85	.0423
1.76	.0142	5.29	.0463
2.21	.0182	5.73	.0503
2.63	.0222	6.17	.0543
3.08	.0262	6.62	.0584
3.53	.0302		

Gerlach, 1859

%	d	%	d
15°			
0	0.99913	7	1.06344
1	1.00823	8	.07282
2	.01733	9	.08231
3	.02647	10	.09281
4	.03560	11	.10150
5	.04484	12	.11120
6	.05408		
0	0.999	7.04	1.063
0.44	1.003	7.48	.068
0.88	1.007	7.92	.072
1.32	.012	8.36	.076
1.76	.015	8.80	.081
2.20	.017	9.24	.085
2.64	.019	9.68	.089
3.08	.027	10.12	.093
3.52	.031	10.56	.097
3.96	.035	11.00	.102
4.40	.039	11.88	.110
4.84	.043	12.32	.115
5.28	.046	12.76	.119
5.72	.051	13.20	.124
6.16	.055		
6.60	.059		

Marignac, 1871

t	d	t	d
3.85 mol%		1 mol%	
11.23	1.23872	0	1.07045
16.30	.23626	9.38	.06849
19.35	.23476	17.06	.06635
23.14	.23284	23.86	.06410
28.73	.22993	29.30	.06209
33.82	.22723	32.26	.06092
39.76	.22393	36.85	.05901
48.20	.21920	40.28	.05751
0	1.24384	0.5 mol%	
9.5	.23948		
13.59	.23752	0	1.03640
23.86	.23241	8.78	.03537
32.26	.22801	13.94	.03435
40.28	.22364	16.84	.03367
1.96 mol%		22.90	.03204
		29.30	.02999
0	1.13350	32.26	.02891
9.38	.13046	36.85	.02715
17.06	.12761		
23.86	.12483	0.25 mol%	
29.30	.12247		
32.26	.12112	0	1.01850
36.85	.11896	8.78	.01792
		13.94	.01713
		16.84	.01656
		22.90	.01510
		27.48	.01379
		32.26	.01223
		36.85	.01059
		40.28	.00925

Favre and Valson, 1874

m	d
24.8°	
0	0.997
0.5	1.059
1.0	1.114
1.5	1.165
2.0	1.213

Kohlrausch, 1879

%	d
18°	
5	1.0450
10	.0915
15	.1426

Volkmann, 1882

%	d
15 -16°	
12.02	1.119
8.55	.0781
3.69	.0329

Nicol, 1885

mol%	d	mol%	d
20°			
0	0.99823	3.09	1.19215
0.50	1.03466	4.25	.25855
0.99	.06744	4.51	.26698
1.81	.11733	5.88	.34014

Le Blanc, 1889

%	d
20°	
0	0.99823
4.76	1.04937
6.46	1.05394
25.51	1.25048

Chernai, 1889

t	d	t	d
sat.sol.			
0	1.040	25	1.209
5	.058	26	.222
10	.078	30	.287
15	.109	33	.312
18	.137	34	.317
20	.156	35	.317

Perkin, 1893

%	d		
	15°	85°	97.9°
26.46	1.2639	1.2211	1.2124
16.01	1.1559	-	-
12.21	1.1141	-	-

Barnes and Scott, 1898

%	d	%	d
17.5°			
0	0.9987	4.015	1.0358
0.2921	1.0014	6.762	.0615
0.5204	.0037	8.544	.0784
1.349	.0109	10.68	.0990
1.818	.0154	11.75	.1094
2.375	.0204	13.06	.1226
2.599	.0225		

Forchheimer, 1900

%	d
20°	
0	0.9982
6.41	1.0570
12.36	1.1138

Dawson and Williams, 1900

%	d	t	%	d	t
sat. sol. 10 aq.					
32.35	1.3218	31.94	33.13	1.3305	33.23
34.26	1.3445	32.93	32.84	1.3257	36.22
33.23	1.3324	32.06	32.20	1.3222	38.31

Brümmer, 1902

%	d	%	d
	15°		
0	0.9993	10.154	1.0931
5.124	1.0462	14.678	1.1393

Berkeley, 1904

%	t	d	%	t	d
4.49	0.70	1.0432	33.06	33.50	1.3307
8.43	10.25	.0802	32.65	38.15	.3229
12.33	15.65	.1150	32.20	44.85	.3136
-	20.35	.1546	31.14	60.10	.2918
21.68	24.90	.2067	30.36	75.05	.2728
25.40	27.65	.2459	29.91	89.85	.2571
29.47	30.20	.2894	29.67	101.9	.2450
32.42	31.95	.3230			

Agerer, 1905

%	d	%	d
	18°		
12.544	1.1167	7.330	0.0664

Dinkhauser, 1905

%	d	%	d
	18°		
1.875	1.0161	7.330	1.0662
4.525	1.0405	12.544	1.1160

Heydweiller, 1921

N	d	N	d
0.2	1.01305	1	1.06186
0.5	1.03166	2	1.1199

Herz and Martin, 1924

t	d	t	d
	11.120g/100cc		
20	1.0925	60	1.0750
30	1.0888	70	1.0695
40	1.0846	80	1.0635
50	1.0800	90	1.0576
	15.373g/100cc		
18	1.1272	32	1.1217
21	1.1263	34	1.1208
24	1.1251	36	1.1199
27	1.1239	38	1.1190
30	1.1227	40	1.1181

Manchot, Jahrstorfer and Zepter, 1924

%	d	
25°		
6.600	1.0550	
13.879	1.1141	
6.640	1.0538	two series
13.730	1.1111	

Gerlach, 1926

%	t	d	%	t	d
1.9	100.2	1.018	7.3	-	1.066
3.2	100.3	.029	10.8	100.8	.101
3.8	100.4	.035	19.3	101.4	.191
6.2	100.6	.056			

Hrynakowski, 1927

sat.sol. 50.20° d= 1.3128

Flöttmann, 1928

%	t	d
11.62	15	1.1083
16.02	20	.1501
21.732	25	.2066

Gibson, 1934

%	π
25°	
0	39.46
5	34.45
10	29.5
15	25.3
20	21.8
22	20.34
24	19.2

%	d	%	d
	25°		
0.9748	1.00587	10.0129	1.08917
1.9656	.01478	10.0245	.08927
2.9148	.02333	12.3968	.11216
4.7746	.04026	14.2624	.13053
4.7721	.04026	21.1287	.20094
4.7725	.04026	21.1293	.20096
9.3124	.08250	21.1362	.20103
	30°		
24.9832	1.23989		
27.9683	.27316		

%	d	%	d
	25°		
0	0.9971	20	1.189
5	1.0423	22	1.210
10	1.089	24	1.231
15	1.138		

Luhdemann, 1935

N	equiv./kg.	d
	25°	
0	0	0.99707
0.8852	0.8955	1.05134
0.9086	0.9195	.05271
1.2264	1.2460	.07138
1.4956	1.4953	.08518
1.5204	1.5510	.08831
1.7430	1.7838	.10100
1.7766	1.8203	.10293

Pearce and Eckstrom, 1937

M	d	M	d
	25°		
0.0	0.997074	0.8	1.090777
0.1	1.009664	1.0	.112294
0.2	.021889	1.5	.163430
0.4	.045658	1.9641	.207404
0.6	.068593		

Usanovich and Mun, 1955

%	d	%	d
	25°		
0	0.9971	12.90	1.1134
3.39	1.0237	16.81	.1565
5.00	.0387	22.66	.2166
8.80	.0766		

Gibson, 1934

t	$Dv/v_{25} \cdot 10^3$			
	0%	1%	5.30%	10%
25.0	0	0	0	0
27.5	0.673	0.709	0.829	0.926
30.0	1.405	1.470	1.701	1.882
32.5	2.194	2.292	2.615	2.867
35.0	3.034	3.162	3.572	3.890
37.5	3.925	4.078	4.564	4.944
	14.81%	19.81%	25%	28%
25.0	0	0	0	0
27.5	0.992	1.044	1.070	1.075
30.0	2.018	2.104	2.158	2.168
32.5	3.054	3.183	3.256	3.282
35.0	4.128	4.294	4.381	4.404
37.5	5.221	5.412	5.506	5.530

Schmidt, 1859

%	t	π
11.65	17.3	35.6
8.47	18.4	38.4
5.34	18.6	40.5
1.5	19.3	44.7
0.6	19.4	45.4

Chernai, 1889

t (sat.sol.)	$\tau \cdot 10^5$		
	20°	35°	50°
-1	20	26	-
"	16	24	-
-0.87	16	22	-
+0.4	-	26	33
1.7	23	25	-
2.3	20	20	33
4.0	26	28	33
9.85	29	31	33
13.78	31	36	33
14.85	-	36	33
15.0	26	37	43
20.5	26	42	-
23.9	26	44	44
23.7	26	-	44
27.6	26	57	45
33.6	26	-	51

Viscosity and surface tension

D'Arcy, 1896

t	η	t	η	t	η
10.44%		11.69%		11.73%	
20.39	1417.5	19.02	1544	19.75	1523
22.22	1354	20.24	1494	21.58	1456
23.49	1315	20.38	1494	23.29	1391
29.00	1159	21.57	1436	24.77	1346
31.32	1103	22.00	1389	30.49	1181
33.50	1054.5	23.37	1388	32.75	1128
35.00	1024	23.39	1448	34.48	1092
40.53	921.9			40.96	960
11.74%		13.77%		15.61%	
19.73	1520	20.0	1654	22.27	1700
20.74	1486	21.39	1597	25.35	1577
21.27	1464	21.47	1591	31.20	1375.7
21.31	1466	22.30	1561	32.33	1345
22.51	1419	23.38	1511	35.54	1256
23.31	1394	23.69	1409	40.16	1146.6
24.76	1349	26.60	1305		
24.78	1349	29.90	1250		
30.2	1186	32.00	1223		
34.38	1086	33.05	1201		
39.70	981	33.97	1147		
		36.13	1052		
		40.52	1508		

Herz and Martin, 1924

t	η	t	η
11.120g/100cc			
20	1410.4	50	772.8
30	1102.3	60	659.1
40	921.6	70	573.4
15.373g/100cc			
18	1721.1	32	1240.6
21	1597.4	34	1187.6
24	1485.9	36	1141.5
27	1386.9	38	1095.2
30	1291.9	40	1005.2

Hrynakowski, 1927

Sat. sol. 50.2° η = 2252

Glass and Madgin, 1934

m	η (water = 1) 40°	38°	36°	34°	33°
0.01783	1.0096	1.0092	1.0088	1.0085	1.0085
0.07156	1.0348	1.0342	1.0335	1.0329	1.0326
0.14437	1.0673	1.0663	1.0654	1.0647	1.0644
0.29296	1.1343	1.1335	1.1328	1.1321	1.1318
0.44938	1.2091	1.2087	1.2081	1.2077	1.2074
0.61247	1.2960	1.2957	1.2955	1.2953	1.2951
0.78607	1.3910	1.3913	1.3917	1.3920	1.3923
0.95279	1.5014	1.5020	1.5026	1.5037	1.5041
1.34412	1.7715	1.7748	1.7776	1.7804	1.7826
1.55108	1.9333	1.9377	1.9423	1.9474	1.9501
1.77155	2.1271	2.1321	2.1386	2.1473	2.1514
2.00111	2.3550	2.3615	2.3730	2.3830	2.3913
2.24276	2.6120	2.6219	2.6337	2.6507	2.6601
2.50123	2.9228	2.9439	2.9653	2.9928	3.0045
3.27607	4.1399	4.1706	4.2850	4.2850	4.3193
	32°	31°	29°	27°	25°
0.01783	1.0083	1.0083	1.0080	1.0079	1.0079
0.07156	1.0323	1.0321	1.0318	1.0316	1.0313
0.14437	1.0641	1.0636	1.0634	1.0630	1.0626
0.29296	1.1315	1.1312	1.1307	1.1301	1.1296
0.44938	1.2073	1.2070	1.2066	1.2064	1.2062
0.61247	1.2950	1.2949	1.2947	1.2945	1.2943
0.78607	1.3925	1.3928	1.3933	1.3938	1.3944
0.95279	1.5047	1.5054	1.5069	1.5086	1.5106
1.34412	1.7848	1.7870	1.8024	1.8024	1.8095
1.55108	1.9527	1.9558	1.9626	1.9691	1.9779
1.77155	2.1568	2.1632	2.1788	2.1945	2.2155
2.00111	2.3985	2.4060	2.4272	2.4498	2.4844
2.24276	2.6710	2.6831	2.7113	2.7459	-
2.50123	3.0202	3.0352	-	-	-
3.01142	3.8391	3.8792	-	-	-

Volkmann, 1882

%	σ
15 - 16°	
12.02	76.2
8.55	75.5
3.69	74.5

Forch, 1899

N	t	σ	$\tau.10^5$
0.1343	17.25	73.14	-
.2689	20.28	73.33	-
.539	20.20	73.67	-
.808	19.78	-	203
.808	14.81	76.96	-
1.078	18.30	74.24	-
.350	19.87	-	195
.350	14.95	74.44	-

Brümmer, 1902

%	σ
15°	
0	74.92
5.124	72.82
10.154	73.36
14.678	72.59

Pann, 1901

%	σ 10°	σ 30°
0	74.009	71.010
4.01	74.656	-
5.07	74.999	-
6.53	75.499	-
11.18	-	73.460
24.8	-	75.871
39.8	-	79.272

Hrynakowski, 1927

Sat. sol. 50.2° σ = 74.60

Glass and Madgin, 1934

m	σ (water = 1)	m	σ (water = 1)
40°			
0.01783	1.000	1.34412	1.058
0.07156	1.003	1.55108	1.064
0.14437	1.007	1.77155	1.073
0.29296	1.015	2.00111	1.082
0.44938	1.022	2.24276	1.093
0.61247	1.030	2.50123	1.102
0.78607	1.038	3.01142	1.128
0.95279	1.044	3.27607	1.140

Optical properties .

Börner, 1869

t	$n_{H\alpha}$	t	$n_{H\beta}$	t	$n_{H\gamma}$
9.9926%					
45.6	1.340748	45.9	1.346785	45.2	1.350117
40.0	.341720	41.0	.347592	41.1	.350869
34.7	.342547	35.4	.348506	35.0	.351834
31.0	.343087	31.2	.349187	31.4	.352960
19.9852%					
47.6	1.351012	49.5	1.356850	48.0	1.360357
42.6	.351870	43.7	.357883	42.5	.361353
37.0	.352835	36.6	.359112	34.85	.362738
32.3	.353621	32.5	.359823	32.8	.363075
27.7	.354335	27.8	.360553	27.85	.363909
29.9848%					
43.3	1.360947	43.7	1.367263	44.4	1.370447
38.8	.361836	39.1	.368113	39.9	.371331
33.9	.362635	33.7	.369068	33.55	.372566
28.8	.363594	28.85	.369881	28.9	.373342

van der Willigen, 1874

spectral lines	n 5,11%	6.80%	8.80%	9.55%	10.51%
21.80°					
A	1.33598	1.33877	1.34153	1.34274	1.34414
a	.33680	.33963	.34241	.34363	.34502
B	.33750	.34032	.34308	.34431	.34568
C	.33824	.34104	.34380	.34503	.34642
D	.34013	.34291	.34571	.34697	.34831
E	.34239	.34515	.34802	.34923	.35063
b	.34282	.34563	.34844	.34968	.35108
F	.34432	.34713	.34996	.35119	.35262
G	.34627	.34912	.35196	.35317	.35465
G	.34783	.35068	.35352	.35477	.35621
H	.34847	.35131	.35419	.35540	.35685
H	.34954	.35240	.35526	.35651	.35792
H	.35076	.35364	.35649	.35773	.35919

Le Blanc, 1889 and Rohland, 1896

%	n_D	%	n_D
20°			
0	1.33395	6.46	1.34243
4.76	.34161	25.51	.37014

Borgesius, 1895

%	t	Dn(sol.-aq.)
4.69	19.2	0.000753
16.11	18.9	0.002931

Jones and Getman, 1904 and Jones, 1904

M	n_D
25°	
0.05	1.32609
.10	.32727
.20	.32949
.50	.33518
1.00	.34389
1.50	.35195
2.00	.35934

Dinkhauser, 1905

%	n_D	%	n_D
18°			
1.875	1.33634	7.330	1.34437
4.525	1.34020	12.544	1.35192

Flöttmann, 1928

%	t	n_D
11.62	15	1.35077
16.02	20	1.35646
21.732	25	1.36380

Luhdemann, 1935

N	equiv./kg	n_D
25°		
0	0	1.33254
0.8852	0.8955	.34136
0.9086	0.9195	.34157
1.0724	1.0873	.34308
1.2264	1.2460	.34447
1.4656	1.4953	.34656
1.5204	1.5510	.34703
1.7430	1.7838	.34892
1.7766	1.8203	.34921

Usanovich and Mun, 1955

%	n_D	%	n_D
25°			
0	(1.332)	12.90	1.3506
3.39	1.3370	16.81	1.3560
5.00	1.3396	22.66	1.3645
8.80	1.3448		

Perkin, 1893

%	(α) magn. (H_2O =1)
15°	
26.46	1.0556
16.01	1.0394
12.21	1.0295
88.9°	
26.16	1.0593

Forchheimer, 1900

%	$(\alpha)_{magn.}$
20°	
6.41	0.375
12.36	0.378

Agerer, 1905

%	t	(α)magn.
12.544	18.2	0.374
7.330	17.9	0.375

Electrical properties

Kohlrausch, 1879

%	κ	τ.10⁴
	18°	
5	406	237
10	684	250
15	883	257

Heim, 1886

t	κ	t	κ
	16.02%		
45	100.95	14	52.18
40	92.40	10	46.72
35	84.53	7	42.85
30	76.53	3	37.96
25	68.83	0	34.50
22	64.29	-2	32.60
20	60.97	-4	29.74
19	59.64	-7	26.64
18	58.09	-8	25.65
16	55.07	-9.5	24.23

Dawson and Williams, 1900

t	κ	t	κ
	sat.sol.		
25.55	1157	33.23	1386
27.00	1214	32.24	1367
27.97	1264	33.23	1405
29.15	1305	33.83	1425
30.98	1349	34.93	1465
31.22	1359	36.87	1531
32.10	1372	39.70	1652

Jones, 1904 and Jones and Getman, 1904

M	λ	M	λ
	0°		
0.05	93.06	1.00	47.13
0.10	84.92	1.50	36.82
0.20	73.09	2.00	28.64
0.50	60.54		

Bachofner, 1904

t	κ 0.5 N	1 N	1.5 N	2 N	2.5 N
10	242	408	537	636	715
15	281	471	624	744	832
20	320	536	709	852	948
25	358	597	797	960	1066
30	396	662	885	1064	1184
35	434	728	974	1170	1305
40	472	796	1064	1278	1428
45	512	867	1158	1392	1554
50	551	938	1252	1504	1680
55	590	1009	1347	1614	1806
60	629	1078	1438	1722	1932
65	669	1151	1532	1834	2058
70	708	1222	1624	1944	2184
75	748	1293	1713	2054	2309
80	786	1364	1804	2162	2436

M	c	κ 20°	40°	60°	80°
0.1	1.422	177	262	346	421
0.2	2.843	274	403	532	664
0.25	3.554	320	472	629	786
0.3	4.265	366	541	725	907
0.4	5.686	454	673	910	1148
0.5	7.108	536	796	1078	1364
0.6	8.530	606	908	1228	1555
0.7	9.951	676	1013	1370	1726
0.75	10.662	709	1064	1438	1804
0.8	11.373	740	1112	1500	1884
0.9	12.794	800	1200	1618	2030
1.0	14.216	852	1278	1722	2162
1.1	15.638	896	1347	1813	2387
1.2	17.059	932	1403	1894	2387
1.25	17.770	948	1428	1932	2436

Heydweiller, 1921

N	λ
	18°
0.2	70.9
0.5	59.8
1	50.7
2	40

Heat constants.

Schüller, 1869

%	U	%	U
9.09	0.9253	20.00	0.8523
13.04	.8959	23.08	.8320
16.67	.8704	28.57	.8074

Marignac, 1876

%	U(21-52°)	%	U(21-52°)
3.79	0.9596	13.62	0.8784
7.29	0.9270	23.98	0.8191

Pagliani, 1880 and 1881

%	U	%	U
	room temp.		
0	1.001		
1.93	0.977	8.97	0.902
3.79	0.959	13.63	0.870
7.31	0.923	19.34	0.843

Gerlach, 1926

%	t	U	%	t	U
1.9	12-15	0.977	-	16-89	0.960
3.2	15-51	.958	7.3	-	.920
-	17-87	.978	10.8	18	.892
3.8	-	.955	19.3	20-23	.843
6.2	14-54	.933	30.3	24-100	.781

Randall and Rossini, 1929

M	U	M	U
	25°		
0.00	0.9979	0.35	0.9460
0.01	.9960	0.50	.9287
0.02	.9942	0.75	.9040
0.05	.9890	1.00	.8835
0.10	.9807	1.25	.8661
0.20	.9657	1.50	.8510

Kobe and Anderson, 1936

t	U	t	U
	sat.sol.		
32.5	0.7946	75-85	0.7734
35-45	.7933	85-95	.7668
45-55	.7893	100	.7596
55-65	.7846	102.8	.7575
65-75	.7793		

Q tr. (10 aq.) = 18.700 cal/mole.

Colson, 1902

Dt on dilution of 100 cc of a solution with 400 cc of water.

t	Dt	t	Dt	t	Dt	t	Dt
29.39%		12.60%		9.80%		9.28%	
10.6	-1.28	12.3	-0.27	12.9	-0.18	12	-0.14
19.0	-1.03	18.1	-0.25	17.1	-0.16	16	-0.125
29.2	-0.80	24	-0.20	29.3	-0.10	40.7	-0.035
34.7	-0.57	44.2	-0.07	39	-0.05	44	-0.025
39.3	-0.46	44.4	-0.065	45	-0.025	46	-0.02
44.6	-0.31	53	-0.03	56	+0.02	54	0
50	-0.20	62.5	+0.03	57.2	+0.02	56	+0.0175
56	-0.07	89	+0.12				
56.7	"	"	"				

Kapustinskii and Ruzavin, 1955

thermic conductivity coefficient $c.10^6$

%	c	%	c
	25°		
6.15	1452	17.6	1445
11.9	1447	18.1	1447

Kapustinskii and Ruzavin, 1956.

m	H
25°	
0.50	0.202
1.00	0.20
1.45	0.20
1.70	0.198

$$H = \frac{(1000 + mM)\,K - 1000\,Ko}{m}$$

m = molality

K = heat conductivity coefficient of the solution

Ko = " " " of water

Water (H_2O) + Sodium ammonium sulfate (NaH_4NO_4S)

Schiff, 1860

%	d	%	d
	15°		
25.19	1.1739	10.08	1.0669
20.15	.1370	5.04	.0328
12.60	.0839		

Water (H_2O) + Sodium acid sulfate ($NaHO_4S$)

Tammann, 1885

%	p	%	p
	100°		
15.41	724.8	47.39	580.3
20.64	708.4	55.16	522.3
25.46	688.8	57.29	504.5
38.83	626.7		

Marignac, 1871

t	d			
	3.85mol%	1.95mol%	1mol%	0.5mol%
0	1.18332	1.10007	1.05267	1.02729
9.62	.17711	.09630	.05046	.02603
12.40	.17537	.09511	.04965	.02547
16.10	.17304	.09348	.04850	.02460
21.90	.16932	.09076	.04647	.02298
27.01	.16602	.08823	.04448	.02131
31.43	.16313	.08597	.04264	.01971
33.80	.16160	.08471	.04160	.01881

Moore, 1895-96

m	d	η
	18°	
0.00	0.9987	1058
0.25	1.0186	1121
0.5	.0386	1164
1.0	.0753	1333
2.0	.1475	1716
4.0	.2810	3041

Hrynakowski, 1927

sat.sol. 50.70° d= 1.5713 η= 6113 σ= 78.75

Gillespie and Wasif, 1953

M	κ	M	κ
	25°		
0.1318	173.6	1.4946	646.8
.3206	299.7	.8502	682.8
.6972	471.1	2.2900	704.2
.9666	550.5		

Water (H_2O) + Sodium selenate (Na_2O_4Se)

Funk, 1900

%	f.t.		%	f.t.	
11.74	0	10 aq.	45.47	35.2	0 aq.
25.01	15		45.26	39.5	
36.91	25.2		44.49	50	
39.18	27		42.83	75	
44.05	30		42.14	100	

sat.sol. 18° d= 1.315

Smits and Mazee, 1928

mol%	f.t.	mol%	f.t.
6.7	76	6.6	269
6.5	86	6.7	274
6.4	107	7.6	333
6.4	232	8.3	372
6.5	262	100	777

tr.t. 10 aq. - 0 aq. = 30.3°

Traube, 1895

%	d	%	d
	15°		
0	0.9991	8.528	1.07834
1.182	1.00982	12.914	1.12220
4.594	1.04109	13.509	1.12827

Water (H_2O) + Sodium thiosulfate ($Na_2S_2O_3$)

Heterogeneous equilibria.

Tammann, 1885

t	p 0%	16.22%	24.49%	33.66%	43.41%
19.69	17.2	-	-	15.0	12.4
30.47	32.7	-	-	28.3	24.0
35.51	43.4	-	-	37.7	32.1
39.26	53.2	50.1	48.3	45.9	39.1
45.19	72.6	69.0	66.8	62.9	54.0
48.67	86.6	82.2	79.3	74.7	65.0
52.47	104.5	99.1	95.5	89.8	77.2
56.27	125.5	119.6	115.0	108.1	93.7
59.42	145.5	138.5	133.4	125.7	109.1
63.31	173.9	164.8	158.5	149.8	129.9
66.78	203.1	193.4	186.2	175.8	153.5
69.46	228.4	217.6	208.8	196.9	172.4
72.64	261.9	248.5	240.2	225.6	197.5
76.49	307.9	293.3	282.2	266.0	233.8
78.64	336.4	319.5	309.9	290.5	255.1
80.96	369.5	352.1	338.9	318.7	280.6
84.74	429.4	409.0	394.3	372.1	328.7
87.50	478.0	455.5	439.0	413.6	365.1
89.53	516.7	491.3	474.3	445.3	394.9
91.72	561.3	533.5	514.5	485.7	429.1
94.38	619.6	587.7	567.0	535.1	473.5
97.32	690.0	654.7	630.9	595.6	521.3
99.97	759.2	723.5	697.8	659.3	585.9

%	p	%	p
	100°		
8.18	744.0	35.28	644.8
12.24	734.0	40.13	613.6
20.16	711.5	45.63	568.9
24.78	695.8	48.36	546.9
29.24	676.6	52.56	509.8
32.20	662.1	64.47	399.0

Speranski, 1911

t	p	t	p
	sat.sol.(5 aq.)		
42.7	14.63	52.15	10.96
46.9	7.75	56.6	14.51
48.3	8.47	58.2	16.84

Perreu, 1930

%	p 20°	p 27°	%	p 20°	p 27°
42.8	15.72	24.09	64.2	-	20.82
50.0	15.07	23.32	66.6	-	20.30
54.5	14.64	22.70	68.7	-	19.80
58.3	13.85	21.97			
61.5	13.55	21.43			

Lescoeur, 1896

t	p 5 aq. + 0 aq.	p sat.sol.
10 E	-	7.0
15	2.8	10.0
20	4.1	12.0
35	11.7	-
40	18.1	33.2

Gerlach, 1886

%	b.t.	%	b.t.
0	100	58.59	114
12.28	101	60.32	115
21.26	102	62.12	116
28.06	103	63.74	117
33.11	104	65.27	118
37.11	105	66.78	119
40.48	106	68.20	120
43.50	107	69.61	121
46.24	108	70.92	122
48.72	109	72.38	123
50.96	110	73.89	124
53.04	111	75.67	125
54.95	112	77.68	126
56.79	113		

Stokes, 1948

m	osmotic coefficient
25°	
0.1	0.805
.2	.774
.3	.753
.4	.741
.5	.731
.6	.724
.7	.719
.8	.713
.9	.709
1.0	.707
.2	.707
.4	.713
.6	.719
.8	.728
2.0	.739
2.5	.779
3.0	.829
3.5	.890

Kremers, 1856

f.t.	%	
0	33.22	
20	40.98	(5+1)
40	51.02	
60	65.79	

Mulder, 1866

%	f.t.	
39.32	16	
42.99	25.25	(5+1)
45.80	32	
53.29	47	

Guthrie, 1878

%	f.t.	%	f.t.
1	-0.1	10	-2.5
2	-0.4	15	-3.9
3	-0.65	20	-5.45
5	-1.2	30	-9.5
6	-1.5	30	-11.0 E

Taylor, 1897-98

%	f.t.		%	f.t.	
34.53	0.2	5aq.	62.92	50	2aq.
38.52	12.2		64.10	52.5	
41.21	19.9		65.33	55.6	
45.85	30		67.40	60	
50.62	40		70.39	72	
54.49	45		71.33	80.5	
63.09	50		71.77	90.5	
			72.67	100	

Young and Burke, 1906

f.t.	%	f.t.	%
	5 aq. (I)		
0	33.40	10	37.37
5	35.33	15	39.11
	2 aq. (I)		
0	52.73	40	59.38
5	53.45	45	60.73
10	53.94	55	63.85
15	54.62	60	65.68
35	58.15	65	68.04
	5 aq. (II)		
0	41.96	10	45.25
5	43.56	15	47.19
	1 aq. (II)		
0	60.47	30	63.56
5	60.74	35	64.32
10	61.04	40	65.22
15	61.57	45	66.02
20	62.11	50	66.82
25	62.73	55	67.90
	6 aq. (III and IV)		
0	46.14	10	51.66
5	48.44	13	54.96
	1 1/2 aq. (III)		
0	57.42	30	60.78
5	57.84	35	61.57
10	58.28	40	62.60
15	58.80	45	63.97
20	59.28	47.5	64.68
25	60.18		
	1 aq. (III)		
47.5	64.78	55	66.45
50	65.30	60	68.07
52.5	65.89		
	1 1/3 aq. (IV)		
0	57.63	30	61.03
5	58.08	35	61.94
10	58.49	40	62.95
15	59.00	45	64.22
20	59.57	50	65.45
25	60.35	55	67.07
	2 aq. (V)		
0	57.63	20	61.02
5	58.23	25	62.30
10	59.05	30	63.56
15	60.02	35	65.27
	1 aq. (V)		
30	63.34	45	65.58
35	64.07	50	66.58
40	64.75	55	67.59
	1/2 aq. (V)		
25	64.21	55	66.57
35	64.60	60	67.40
40	64.99	65	68.24
45	65.61	70	69.06
50	66.02		
	0 aq.		
40	67.40	65	68.80
45	67.60	70	69.05
50	67.76	75	69.35
55	68.15	80	69.86
60	68.48		

48.17°	5 aq. (I) - 2 aq.
40.65°	4 aq. (II) - 1 aq. (II)
31.50°	4 aq. (II) - 2 aq. (I)
30.46°	5 aq. (II) - 2 aq. (I)
30.22°	5 aq. (II) - 4 aq. (II)
14.35°	6 aq. (III) - 6 aq. (IV)
48.45°	5 aq. (I)
41.65°	4 aq. (II)

Picon, 1924

mol%	f.t.	mol%	f.t.
0.85	-1.8 ice	4.24	-9.1
1.39	-2.4	4.39	-9.6
1.74	-3.5	4.5	-10
2.09	-4.3	4.71	-10.6
2.57	-5	5.21	-11
2.63	-5.8	5.7	-14
3.4	-6.8	5.73	-14
3.53	-7.4	6.02	-15.2
3.57	-7.4	6.24	-16
3.74	-7.6	6.56	-17.5
3.94	-8.4	7	-20.5
4.14	-8.6	7.25	-22.5

mol %	f.t.	mol %	f.t.
	12 aq.		
7.49	-16.5	8.02	-16.3
	10 aq.(I)		
6.01	-13.4	7.25	- 9.
6.34	-11.5	7.49	- 8.6
6.56	-10.5	8.02	- 8.2
7	- 9.5	8.45	- 7.9
	6 aq.(II)		
6.24	-16	9	+ 3.4
6.56	-13	10	8
7	- 9.5	11.92	12.4
8.02	- 3	13	14
8.45	- 0.5		
	5 aq.(II)		
6.34	-14	10	20.8
6.56	- 9	10.98	25.5
7.03	- 2	11.92	28.1
7.56	+ 2.5	13.01	30.1
7.98	5.4		
	5 aq. (I)		
4.92	- 4.8	8.54	31
4.94	- 3.4	9.68	37.6
5.21	- 1	10.97	41.5
5.66	+ 3.6	12.50	45.5
5.73	4.5	13.46	46.9
6.02	8.5	14.28	47.8
6.23	9.75	15.02	48
6.43	12.7	15.72	48.2
6.87	15.7	16.55	48.3
7.01	18.4	17.35	48.6
7.64	23.1	17.80	48
8.02	25.8	18.15	48
	4 aq.(II)		
14.23	34.9	16.74	40
15.00	37.9	17.23	40.1
16.05	39		
	2 aq.(I)		
15.57	49.6	18.23	61.3
16.81	56.7	18.30	62.3
17.40	59	19.20	64.9
19.20	65	19.68	66.5
19.40	65.6	19.71	66.8
19.51	66		
	1 aq.(II)		
17.06	46.2	20.19	63.5
18.08	51.9	20.48	65.4
18.10	52.5	20.79	66.5
19.20	59	20.86	66.8
19.34	59.5	21.17	67.5
19.75	61.4	21.39	69.1
19.84	62.4	21.64	69.5
	1 aq.(I)		
17.92	45.6	20.22	61.7
18.49	49.5	20.64	63.75
18.62	51.5	20.84	65
19.05	53.5	21.86	70
19.33	55.5	22.17	71.5
19.61	57.5	22.26	72
19.84	59	22.45	72.5
20.06	60.5		
	1/2 aq.(I)		
18.77	59.5	19.80	70
19.06	63.5	20.07	72
19.32	65	20.44	75
19.40	66.5	20.72	76
19.44	67.2	22.26	78
19.75	69.5		
	0 aq.		
19.69	61	20.80	83
19.72	61.7	20.97	88
19.90	67	21.19	91
19.94	68	21.30	93.5
19.96	68.5	21.58	97
19.98	69.5	22.32	107
20.08	69.8	22.60	112
20.22	74	22.69	112.5
20.42	76.5	22.91	114.5
20.48	79.75		

Benrath, 1942

%	f.t.	%	f.t.
71.4	111	74.6	148
72.4	125	76.3	159
72.8	130	77.7	171
73.4	136	79.1	179

Mathieu, 1949

$f.t.(0-48°) = -169.512 + 6.90117p - 0.0546788p^2$
where p= wt%

%	f.t. 1	250	500	750	1000
	atm.				
43.49	27.2	27.61	28.17	28.88	29.75
45.43	31.15	31.59	32.20	32.9	33.79
48.765	37	37.8	38.6	39.4	40.20
56.61	41	41.54	42.35	43.42	44.76
52.76	42.4	43.04	43.92	45.03	46.38
55.33	45	46.53	47.76	48.72	49.38
58.715	47.2	49.03	50.51	51.67	52.50

Jones B.M., 1909

%	t.spont.cryst.	%	t.spont.cryst.
0	-0.7 ice	9.0 laq.	64.16
3.86	-1.9	18.9	65.57
6.65	-2.6	21.9	65.71
12.99	-4.3	28.0	66.79
20.92	-6.9	31.7	67.31
22.33	-7.2	39.2	69.02
22.46	-7.6	45.3	70.09
26.10	-9.2	52.9	71.67
		53.7	71.85

Properties of phases.

Schiff, 1860

%	d	%	d
	19°		
0.64	1.0036	16.55	1.1422
1.27	.0089	17.19	.1481
1.91	.0172	17.82	.1540
2.55	.0195	18.46	.1599
3.19	.0248	19.11	.1658
3.82	.0301	19.75	.1720
4.46	.0354	20.37	.1781
5.10	.0407	21.00	.1843
5.73	.0460	21.64	.1905
6.37	.0512	22.27	.1967
7.00	.0567	22.91	.2027
7.64	.0622	23.55	.2091
8.27	.0678	24.19	.2153
8.91	.0734	24.83	.2215
9.55	.0790	25.48	.2278
10.19	.0847	26.12	.2343
10.82	.0902	26.75	.2407
11.46	.0958	27.39	.2472
12.10	.1014	28.02	.2538
12.74	.1070	28.66	.2604
13.36	.1128	29.29	.2670
14.00	.1186	29.93	.2736
14.64	.1245	30.57	.2802
15.27	.1304	31.21	.2868
15.91	.1363	31.86	.2934

Damien, 1881

%	d	%	d
	19°		
0	0.99847	25.77	1.2219
6.88	1.0553	33.631	.3012
8.03	.0667	42.24	.4056
10.03	.0830	53.03	.5247
15.36	.1298	57.54	.5733
20.84	.1785	100	.6371

Nicol, 1885

mol%	d	mol%	d
	20°		
7.12	1.38896	10.45	1.50136
8.44	.43231	10.82	.50383
9.14	.46231	13.15	.59006
9.97	.49392	16.67	.67335

Schiff and Monsacchi, 1897

%	d
19°	
0	0.9984
36.24	1.3434
40.46	1.3875

Schiff and Monsacchi, 1896 and 1897

%	d	%	d
	19°		
0	0.9984	38.226	1.3629
3.185	1.0264	44.597	.4341
6.371	.0529	47.782	.4708
12.742	.1087	50.968	.5083
19.113	.1676	57.339	.5856
25.484	.2297	63.710	.6659
31.855	.2954		

Chèneveau, 1907

%	d	%	d
	15°		
0	0.9991	23.85	1.2128
4.63	1.0383	27.10	1.2454
8.93	1.0750	30.10	1.2758
12.98	1.1117	33.03	1.3069
16.82	1.1464	35.78	1.3367
20.45	1.1810		

Rakshit, 1925

%	d	%	d
	20°		
0.64	1.00407	19.1	1.14292
3.2	1.02473	31.9	1.23041
6.4	1.04974		

Bak, 1939 (fig.)

%	η (arbitrary units)			
	40°	60°	70°	80°
0	8	6	4	2
20	10.5	7.5	5.5	3
40	25	18	12	9
52	64	42	35	25
65	-	110	75	58
70	-	125	95	65

Lecoq de Boisbaudran, 1895

%	Dv	%	Dv
	15.1°		
0.96	-17.97	38.2	-2.74
4.5	-15.09	54.2	+2.00
7.7	-13.35	100.0	+4.54
22.3	- 7.68	$Na_2S_2O_3$ + 5 aq.	

Dv in % of crystal volume

d of crystal at 15° = 1.752

Taimni, 1929

%	η		
	45°	40°	35°
42.2	3100	3500	4000
44.8	3800	4400	5100
47.9	4900	5600	6600
50.7	5900	6800	8100
54.0	8800	10300	12400
	30°	25°	20°
42.2	4700	5600	6800
44.8	6000	7100	8700
47.9	7900	9600	11900
50.7	9800	12200	15200
54.0	15200	19300	25200

Morgan and Schramm, 1913

%	σ	%	σ
	40°		
0	69.33	35.11	79.19
7.21	70.65	40.46	82.02
13.70	71.83	44.33	84.88
19.98	73.53	46.71	86.22
23.27	74.50	49.58	88.02
27.67	75.94	56.87	92.27
29.87	76.75	60.60	95.38
32.45	77.93	63.88	97.08

Damien, 1881

%	n		
	H_α	H_β	H_γ
		19°	
0	1.33117	1.33716	1.34045
6.88	.34581	.35227	.35586
8.03	.34905	.35560	.35941
10.03	.35239	.35871	.36239
15.36	.36530	.37280	.37697
20.84	.37478	.38024	.38348
25.77	.38769	.39605	.40074
33.631	.40689	.41602	.42112
42.261	.42786	.43558	.43971
53.03	.45304	.45062	.46530
57.54	.46224	.47051	.47530
100	.48071	.49184	.49795

Chèneveau, 1907

%	n_D	%	n_D
	15°		
0	1.3334	23.85	1.3855
4.63	.3432	27.10	.3933
8.93	.3521	30.10	.4004
12.98	.3610	33.03	.4075
16.82	.3694	35.78	.4142
20.45	.3778		

Blythwood and Marchant, 1899

M	th
1	10.6
2	8.5
5	7.5

th = thickness of the solution in mm producing the same X-ray absorption as 20 mm of water .

Küster and Thiel, 1899

mol %	κ (KCl = 1)	mol %	κ (KCl = 1)
	60°		
13.07	930	15.15	796
13.53	904	15.60	777
13.79	882	16.10	716
14.02	865	16.86	670
14.72	829	17.57	635

Bak, 1939 (fig.)

%	κ(arbitrary units)			
	40°	60°	70°	80°
5	7	8.5	9	9.5
20	22	28	31.5	35
30	22	28	32.5	36
40	20	26	31	35
60	3	9	13	20
62	-	-	-	16
70	-	12	15	17

Water (H_2O) + Sodium hyposulfite ($Na_2S_2O_4$)

Jellinek, 1911

%	f.t.	%	f.t.
0 ice	0	8.45	-2.21
1.04	-0.28	10.09	-2.66
2.06	-0.56	11.77	-3.15
3.09	-0.84	13.37	-3.63
-	-1.10	15.01	-4.17
4.80	-1.30	16.16	-4.48
5.93	-1.56	15.97	-4.58
6.55	-1.76		
+20	17.89% 2 aq.		
	19.42% 0 aq.		

Water (H_2O) + Sodium dithionate ($Na_2S_2O_6$)

Tammann, 1885

%	p	%	p
	100°		
13.06	739.7	34.81	675.3
24.77	710.4	37.94	661.3
27.08	703.6		

Ishikawa and Oku, 1932

%	f.t.		d
1.45	- 0.66	ice	
3.64	- 0.+85		
5.13	- 1.082		
5.00	- 1.136 E		
5.60	- 0.73	8 aq.	
5.69	- 0.49		
5.86	- 0.14		
5.90	- 0.00		
5.90	0	6 aq.	-
6.04	0.5		-
6.21	1		-
6.89	3		1.054
7.81	5		1.056
8.95	7.5		1.068
9.78	9		-
10.18	10.		-
7.27	0		1.0565 2 aq.
8.54	5		-
9.40	8		-
10.00	10		1.0773
13.13	20		1.0984
16.42	30		1.1216
19.79	40		1.1460
23.18	50		1.1705
26.50	60		1.1920
29.55	70		1.2191
33.00	80		1.2425
36.02	90		1.2711
39.29	100		1.2981

De Baat, 1926

%	f.t.	%	f.t.
6.5	0	13.39	20
10.63	12	17.32	30

Water (H_2O) + Sodium trithionate ($Na_2S_3O_6$)

Kurtenacker and Laszlo, 1938

%	f.t.	
37.7	0	
52.5	20	3 aq.
58.9	30	

Water (H_2O) + Sodium tetrathionate ($Na_2S_4O_6$)

Kurtenacker and Laszlo, 1938

%	f.t.	
41.3	0	2 aq.
49.9	20	
54.9	30	

Water (H_2O) + Sodium pentathionate ($Na_2S_5O_6$)

Kurtenacker and Laszlo, 1938

%	f.t.
32.9	0
52.0	20

Water (H_2O) + Sodium formate ($NaCHO_2$)

Tammann, 1885

%	p	%	p
	100°		
9.63	721.6	44.73	487.2
16.52	685.4	47.44	468.0
24.74	636.2	53.35	420.3
33.22	576.4	54.61	408.8
36.14	527.4	60.01	373.0
41.50	512.3		

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.6712	0.6486	1.877	1.756
.6872	.6654	.953	.852
.6994	.6773	2.402	2.266
.7126	.6893	.500	.363
.7285	.7030	.534	.387
.7510	.7229	.602	.452
.9950	.9540	.699	.540
1.096	1.050	.742	.583
.146	.097	.930	.755
.253	.200	.920	.760
.317	.258	3.057	.876
.474	.405	.128	.910
.754	.668	.123	.945
.812	.716	.153	.962
.873	.772	.585	.373

m_1 = molality of KCl m_2 = molality of $NaCHO_2$

Groschuff, 1903

%	f.t.		%	f.t.	
22.80	-20	3 aq.	44.73	18	2 aq.
30.47	0		46.86	18	
41.88	+15		48.22	21	
44.92	18		50.53	23	
49.22	18	0 aq.	56.82	74.5	0 aq.
49.95	25.5		61.54	100.5	
50.44	29		66.20	123	
53.80	54				

Sidgwick and Gentle, 1922

%	f.t.		%	f.t.
7.24	-4.29	ice	-16.95	11.02
11.42	-7.02		-24.86	18.04

Elöd and Tremmel, 1927

%	f.t.		%	f.t.	
39.8	13.0	3 aq.	42.7	16.0	2 aq.
41.1	15.0		43.5	18.0	
50.6	29.0	0 aq.			
52.7	45.0				

Kannonikov, 1885

%	t	d
0	26.6	0.9966
17.14	26.6	1.1135

Perkin, 1891

%	d 15°	d 25°
0	0.9991	0.9971
7.39	1.1541	1.1492
15.63	1.2894	1.2831

Wasastjerna, 1920

M	d	M	d
	18°		25°
0.0000	0.99862	0.0000	0.99707
0.6562	1.02645	0.6549	1.02447
1.6032	.06476	1.5994	.06222
2.7264	.10866	2.7189	.10560
3.3328	.13141	3.3235	.12824
4.7202	.18246	4.7054	.17875
5.2082	.19984	5.1914	.19596
5.9112	.22494	5.8918	.22092
6.8774	.25786	6.8538	.25354
7.2718	.27117	7.2465	.26674

de Garcia, 1920

N	d	N	d
	20.5°		
4	1.148960	0.25	1.009238
2	.076866	0.125	.004426
1	.038732	0.062	.001641
0.5	.019210		

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.02114	3	1.1196
1	.04173	4	.1565
2	.0816		

Lühdemann, 1935

N	d	N	d
	25°		
0	0.99707	3.1822	1.12305
0.5350	1.01951	3.8716	.14866
1.1132	.04303	4.5206	.17226
1.8444	.07197	5.2730	.19914
2.4928	.09702		

Kanonnikoff, 1885

%	t	n H_α	n D	n H_β
0	26.6	1.33045	1.33233	1.33642
17.14	26.6	1.352125	1.354062	1.358875

Wasastjerna, 1920

M	n 6563 Å	n 5893	n 4861
		18°	
0.0000	1.33168	1.33348	1.33764
0.6562	.33720	.33911	.34339
1.6032	.34463	.34668	.35109
2.7264	.35281	.35487	.35956
3.3328	.35697	.36480	.36807
4.7202	.36589	.36807	.37295
5.2082	.36893	.37108	.37593
5.9112	.37305	.37526	.38014
6.8774	.37829	.38049	.38573
7.2718	.38039	.38271	.38789
		25°	
0.0000	1.33108	1.33348	1.33701
0.6549	.33641	.33911	.34255
1.5994	.34378	.34668	.35020
2.7189	.35184	.35487	.35850
3.3235	.35598	.35901	.36283
4.7054	.36480	.36807	.37182
5.1914	.36774	.37108	.37480
5.8918	.37175	.37526	.37889
6.8538	.37712	.38049	.38454
7.2465	.37919	.38271	.38661

de Garcia, 1920

N	n_D	N	n_D
	20.5°		
4	1.3593	0.25	1.3345
2	.3470	0.125	.3337
1	.3401	0.062	.3334
0.5	.3365		

Lühdemann, 1935

N	n_D	N	n_D
	25°		
0	1.33254	3.1822	1.35667
0.5350	1.33707	3.8716	1.36127
1.1132	1.34165°	4.5206	1.36540
1.8444	1.34720	5.2730	1.36998
2.4928	1.35190		

Perkin, 1891

%	t	$(\alpha)_D^{magn.}$ (H_2O =1)
7.39	16.8	1.0525
15.63	18.75	1.0904

Heydweiller, 1921

N	λ
18°	
0.5	61.4
1	53.7
2	43.08
3	34.8
4	28.15

Jones, 1904 and Jones and Bassett, 1904 and 1905

M	λ	M	λ
	0°		
0.10	97.00	0.50	72.50
0.20	89.40	0.60	69.00
0.30	80.00	0.80	57.59
0.40	76.00	1.00	53.95

Marignac, 1876

%	U
(21-52°)	
0	1
4.32	0.9511
8.28	0.9134
15.30	0.8560
26.54	0.7810

Water (H_2O) + Sodium acetate ($NaC_2H_3O_2$)

Heterogeneous equilibria.

Lescoeur, 1885 and 1890

%	p	%	p
	20°		
-	17.54	35.67	12.04
4.82	16.94	39.99	11.04
9.83	15.54	47.41	8.84
19.80	15.04	50.94	7.44
29.79	13.34	54.88	5.84

Tammann, 1885

%	p	%	p
	100°		
7.13	737.7	28.77	618.9
10.92	721.8	30.31	606.8
14.47	706.1	32.81	588.0
19.21	680.3	38.96	538.5
21.56	666.8	41.78	515.6
23.81	653.5	43.82	498.9
26.90	633.2		

Lescoeur, 1885 and 1890

t	p			
	6aq.-3aq.	sat.sol.	3aq.-0aq.	sat.sol.
	43.15 %		60.29 %	
10	-	5.6	-	3.2
15	2.4	-	2.05	4.45
20	4.4	11.3	3.6	6.4
25	6.9	-	-	9.2
30	10.5	20.4	6.9	12.9
35	15.3	25.4	10.5	15.9
40	20.6	31.2	-	20.9
50	35.3	-	-	-
60	-	-	-	67
80	-	-	-	162

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.940	1.0	1.002
0.2	0.939	1.2	1.018
0.3	0.945	1.4	1.038
0.4	0.951	1.6	1.057
0.5	0.959	1.8	1.074
0.6	0.967	2.0	1.092
0.7	0.977	2.5=	1.137
0.8	0.986	3.0	1.181
0.9	0.994	3.5	1.223

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1810	0.4580	2.337	2.025
.6292	.5886	.405	.023
.6872	.6378	.534	.128
.8530	.7850	.602	.183
.9950	.9010	.699	.255
1.065	.9610	.742	.290
.110	1.013	.911	.420
.123	.088	.930	.431
.214	.139	3.057	.528
.278	.238	.098	.557
.398	.440	.587	.910
.647	.440	.585	.922
.722	.500	.685	.997
.763	.534		

m_1 = molality of potassium chloride

m_2 = " of sodium acetate

Legrand, 1835

%	b.t.	%	b.t.
0	100	47.61	113
9.01	101	49.70	114
14.96	102	51.71	115
19.42	103	53.66	116
23.38	104	55.58	117
26.85	105	57.42	118
30.02	106	59.21	119
33.02	107	60.95	120
35.82	108	62.60	121
38.42	109	64.20	122
40.90	110	65.71	123
43.24	111	67.15	124
45.47	112	67.63	124.37

Gerlach, 1886

%	b.t.	%	b.t.
0	100	45.95	113
7.84	101	47.92	114
13.79	102	49.88	115
19.03	103	51.80	116
23.38	104	53.70	117
26.74	105	55.75	118
29.83	106	57.63	119
32.66	107	59.34	120
35.06	108	60.31	120.54
37.50	109	61.08	121
39.76	110	62.68	122
41.86	111	64.41	123
43.96	112	65.98	124
		67.43	125

Mulder, 1866

%	f.t.	%	f.t.
32.48	16	41.59	47
34.18	25.25	45.59	52
36.02	32		

Rüdorff, 1872

%	f.t.	%	f.t.
3.42	-1.55	11.67	-6.60
4.47	-2.10	13.92	-8.30
5.48	-2.60	15.96	-10.20
7.40	-3.75	17.23	-11.20
10.05	-5.45	17.84	-12.00

Guthrie, 1876

%	f.t.		%	f.t.	
5	-2.2	ice	22	-16.0	
10	-5.1		23.3	-18.0	E
15	-9.1		26.6	0.0	x aq.
20	-14.0				

Schiavon, 1902

%	f.t.		%	f.t.
28.57	9	+3 aq.	37.88	37
33.33	13		40.49	41

Green, 1908

%	f.t.		%	f.t.	
54.35	0	0 aq.	59.84	72.5	
55.56	25		60.68	83	
56.34	34.8		60.68	83.5	
56.70	43		62.96	100	
57.18	50		65.88	123	
58.42	61				
			46.66	51.5	3 aq.
			50.48	55.5	

Sidgwick and Gentle, 1922

%	f.t.		%	f.t.
6.107	-2.96	ice	16.62	-11.12
11.27	-6.42		20.57	-15.17

Properties of phases.

Gerlach, 1869

%	d	%	d
17.5°			
0	0.9987	20	1.062
5	1.014	25	1.0781
10	1.030	sat.sol.	1.169
15	1.046		

Franz, 1872

%	d	%	d
17.5°			
0	0.9987	20	1.1060
5	1.0279	25	1.1360
10	1.0524	30	1.1691
15	1.0788	sat.sol.	1.1826

Favre and Valson, 1874

m	d	m	d
23.25°			
0	0.9975	2.0	1.128
0.5	1.039	2.5	1.153
1.0	1.072	3.0	1.175
1.5	1.101		

Kohlrausch, 1879

%	d	%	d
18°			
5	1.025	30	1.159
10	1.051	32	1.170
20	1.104		

Gerlach, 1887

%	d	%	d	%	d
17.5°					
0	0.999	11	1.056	22	1.128
1	1.004	12	1.062	23	1.123
2	1.009	13	1.067	24	1.129
3	1.015	14	1.073	25	1.135
4	1.020	15	1.078	26	1.141
5	1.025	16	1.084	27	1.147
6	1.030	17	1.089	28	1.152
7	1.035	18	1.095	29	1.158
8	1.041	19	1.100	30	1.164
9	1.046	20	1.106	31	1.170
10	1.051	21	1.112	(sat.sol.)	

Le Blanc, 1889

%	d	%	d
	20°		
0	0.99823	21.81	1.11439
5.41	1.02634	45.89	1.10222
9.70	1.04878		

Bindel, 1890

%	d	%	d
	20°		
8.35	1.0442	23.30	1.1316
18.55	1.0985	31.30	1.1718

Perkin, 1891

mol%	d 15°	d 25°
0	0.9991	0.9971
8.18	1.1568	1.1519

8.18 mol% 15.2° $(\alpha)_D^{magn.} = 1.0639$

Rubien, 1911 and Heydweiller, 1912

M	d	M	d
	18°		
0.1013	1.00274	2.039	1.07892
0.5112	.01978	4.044	1.15180
1.020	.03987		

Wasastjerna, 1920

M	d	M	d
	18°		
0.0000	0.99862	1.5624	1.06231
0.2707	1.01016	1.7759	1.07057
0.5413	1.02136	1.9893	1.07876
0.7834	1.03126	2.3540	1.09269
1.0254	1.04103	2.7184	1.10644
1.2939	1.05169		
	24.25°		
0.0000	0.99726	1.5593	1.06021
0.2703	1.00866	1.7722	1.06833
0.5404	1.01972	1.9851	1.07650
0.7821	1.02951	2.3486	1.09019
1.0236	1.03917	2.7122	1.10393
1.2914	1.04968		

de Garcia, 1920

M	d	M	d
	20°		
0.062	1.001248	0.5	1.018063
0.125	.003826	1	.036749
0.25	.008650	2	.072469
		4	.140618

Rakshit, 1925

%	d	%	d
	20°		
0.6	1.00192	18	1.08597
3	1.01400	30	1.13977
6	1.02912		

Guillaume, 1946

%	d
	20°
10.7	1.0553
20.0	1.1072

de Heen, 1881

t	vol	t	vol
	25.90 %		
10.00	1.000000	41.80	1.013791
15.20	.001952	49.89	.017939
20.32	.003992	55.96	.021117
27.50	.006835	62.30	.024658
34.28	.010071		
	19.25 %		
10.00	1.000000	40.92	1.011984
15.40	.001792	48.08	.015400
21.12	.003741	55.36	.019109
28.74	.006710	61.61	.022517
35.80	.009682	68.50	.026437
	13.00 %		
10.00	1.000000	40.51	1.010482
14.90	.001313	47.63	.013699
20.32	.002917	55.54	.017595
28.24	.005574	61.83	.020858
35.68	.008482	68.43	.024608

Walter, 1889

%	n_D	%	n_D
	15°		
0	1.3334	12.5	1.3506
6.0	.3417	22.0	.3648
9.6	.3468	27.6	.3731

Le Blanc, 1889

%	n_D	%	n_D
	20°		
0	1.33325	21.81	1.36371
5.41	.34085	45.89	.36048
9.70	.34671		

Rubien, 1911

M	n_D	M	n_D
	18°		
0	1.33327	1.020	1.34436
0.1013	.33443	2.039	.35463
0.5112	.33895	4.044	.37337

Heydweiller, 1913

M	n_D	M	n_D
	18°		
0	1.33327	1.0	1.34414
0.1	.3342	2.0	.35423
0.5	.33883	4.0	.37298

de Garcia, 1920

M	n_D	M	n_D
	20°		
0.062	1.3333	1	1.3428
0.125	.3340	2	.3526
0.25	.3353	4	.3706
0.5	.3378		

Wasastjerna, 1920

M	n 6563 Å	5893	4861
		18°	
0.0000	1.33168	1.33348	1.33764
.2707	.33475	.33656	.34081
.5413	.33766	.33954	.34385
.7834	.34032	.34222	.34650
1.0254	.34293	.34476	.34916
.2939	.34575	.34759	.35210
.5624	.34850	.35046	.35497
.7759	.35058	.35259	.35717
.9893	.35278	.35479	.35935
2.3540	.35640	.35835	.36306
.7184	.35991	.36201	.36668
		24.25°	
0.0000	1.33116	1.33299	1.33711
0.2703	.33420	.33605	.34018
.5404	.33709	.33892	.34323
.7821	.33963	.34149	.34583
1.0236	.34221	.34415	.34887
.2914	.34493	.34693	.35125
.5593	.34743	.34970	.35414
.7722	.34980	.35183	.35629
.9851	.35197	.35397	.35846
2.3486	.35562	.35761	.36220
.7122	.35907	.36105	.36575

Guillaume, 1946

%	n_{5780}
20°	
10.7	1.3486
20.0	1.3627

Kohlrausch, 1879

%	$\varkappa$	$\tau.10^4$
	18°	
5	293	252
10	478	260
20	647	295
30	597	352
32	567	373

Heydweiller, 1912

M	$\varkappa$	M	$\varkappa$
	18°		
0.1013	2.0	2.039	604.7
0.5112	250.5	4.044	621.1
1.020	418.3		

Tammann and Rohann, 1929

P kg	t	Dλ 0.00063 N	0.001 N	0.01 N
500	20	3.73	3.39	2.99
	40	1.71	1.50	1.41
1000	20	6.28	5.66	5.09
	40	2.56	2.31	2.16
1500	20	7.82	6.99	6.10
	40	2.85	2.50	2.26
2000	20	8.62	7.77	6.31
	40	2.83	2.18	1.89
2500	20	8.52	7.63	5.99
	40	2.24	1.47	1.18
3000	20	7.76	6.88	5.26
	40	1.33	0.59	0.27
		0.1 N	0.5 N	1.26 N
500	20	2.65	2.49	1.71
	40	1.29	1.27	1.10
1000	20	4.50	4.21	2.73
	40	2.03	1.92	1.38
1500	20	5.46	5.5	3.19
	40	2.08	1.93	1.12
2000	20	5.64	5.21	2.99
	40	1.68	1.47	0.61
2500	20	5.23	4.79	2.34
	40	0.93	0.75	0.05
3000	20	4.47	3.56	1.12
	40	0.00	-0.25	-0.87

$D_\lambda = (\lambda_P - \lambda_{P=1} / \lambda_{P=1}).\ 100$

Heat constants

Bindel, 1890

%	U	Q diss.
	20°	
8.35	0.938	-4794
18.55	.870	-4926
23.30	.855	-4968
31.30	.818	-5222

Pagliani, 1880 and 1881

%	U
18°	
4.35	0.974
8.35	.942
15.41	.895
23.29	.871

Richards and Gucker, 1925

t	U
3.85 mol%	
16	0.90003
18	.90102
20	.90212

Water (H_2O) + Sodium propionate ($NaC_3H_5O_2$)

Tammann, 1885

%	p	%	p
	100°		
10.20	733.6	50.67	512.9
19.38	700.7	55.35	480.9
34.67	623.3	59.53	449.6

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.6547	0.5972	1.962	1.622
.6652	.6053	2.100	.715
.6995	.6345	.127	.738
.7218	.6518	.532	2.031
.7555	.6817	.534	.038
.9950	.8780	.602	.086
1.064	.9350	.699	.157
.123	.9828	.742	.187
.146	.9970	.930	.320
.214	1.054	3.057	.415
.398	.198	.098	.442
.545	.305	.153	.478
.677	.408	.585	.784
.860	.546		

m_1= molality of potassium chloride

m_2= " " sodium propionate

Perkin, 1891

%	d 15°	25°	(α) magn. 18.2°
43.24	1.1948	1.1878	1.0938
0	0.9991	0.9971	

Heydweiller, 1921

M	d	M	d
	18°		
0.2	1.00842	2	1.0802
0.5	1.02086	3	1.1174
1	1.04118	4	1.1571

de Garcia, 1920

M	d	n_D	M	d	n_D
		20°			
0.062	1.001078	1.3335	1	1.034663	1.3455
0.125	.003369	.3343	2	.068930	.3580
0.25	.007898	.3360	4	.132369	.3805
0.5	.016963	.3392			

Water (H_2O) + Sodium butyrate ($NaC_4H_7O_2$)

Tammann, 1885

%	p	%	p
100°			
10.69	731.9	38.57	612.6
16.95	710.2	43.62	584.9
29.85	655.8	56.96	492.5

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.8090	0.7060	2.751	2.122
0.8268	0.7172	3.069	2.355
0.8880	0.7680	3.380	2.580
0.9253	0.7968	3.974	3.023
1.416	1.164	4.119	3.130
1.899	1.441	4.568	3.478
2.050	1.610	4.560	3.478
2.531	1.960		

m_1 = m of KCl m_2 = m of $NaC_4H_7O_2$

Perkin, 1891

%	d 15°	d 25°
47.29	1.1678	1.1612
0	0.9991	0.9971

47.29% 18.5° $(\alpha)^{magn.}$ = 1.0968

Vorländer and Kirchner, 1930 - 1931

47 % 20° d = 1.16 η (water=1) = 26.2

Specific mecanical birefringence (anormal) = -0.13

de Garcia, 1920

M	d	n_D
	20.5°	
0.062	1.001471	1.3339
.125	.003872	.3350
.25	.008096	.3372
.5	.016970	.3410
1	.034095	.3485
2	.066991	.3633
4	.120180	.3889

Water (H_2O) + Sodium isobutyrate ($NaC_4H_7O_2$)

Tammann, 1885

%	p	%	p
100°			
12.07	728.7	40.81	595.5
16.40	712.9	45.44	567.5
29.20	657.4	58.26	474.8

de Garcia, 1920

M	d	n_D
	22.5°	
0.062	1.001306	1.3338
.125	.003909	.3348
.25	.009080	.3371
.5	.018862	.3410
1	.038314	.3492
2	.076586	.3650
4	.138816	.3916

Water (H_2O) + Sodium valerate ($NaC_5H_9O_2$)

Tammann, 1885

%	p	%	p
100°			
9.42	740.3	37.69	647.4
17.92	715.6	48.88	600.8
32.98	664.6	59.32	530.5

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.6530	0.5825	1.877	1.513
0.8760	0.7250	2.100	1.684
0.995	0.853	2.500	1.998
0.9978	0.855	2.602	2.083
1.065	0.906	2.699	2.175
1.101	0.934	2.742	2.231
1.123	0.937	2.930	2.424
1.219	0.953	3.057	2.582
1.263	1.024	3.098	2.613
1.322	1.053	3.153	2.703
1.398	1.157	3.417	3.078
1.595	1.266	3.345	3.139
1.727	1.401	3.456	3.281
1.750	1.423	3.515	3.373

m_1 = m of KCl m_2 = m of $NaC_5H_9O_2$

de Garcia, 1920

M	d	n_D
	22.5°	
0.062	1.000921	1.3339
0.125	.003040	.3351
0.25	.007310	.3373
0.5	.016176	.3417
1	.033068	.3506
2	.063190	.3677
4	.105240	.4000

Water + Sodium caproate ($NaC_6H_{11}O_2$)

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.2795	0.2637	1.678	1.495
0.3045	0.2861	1.721	1.538
0.3276	0.3061	1.754	1.584
0.3526	0.3291	1.883	1.758
0.4044	0.3734	1.959	1.774
0.5592	0.5030	2.205	2.418
0.6533	0.5048	2.253	2.512
0.6861	0.5795	2.307	2.620
0.6928	0.6102	2.305	2.651
0.7510	0.6223	2.500	3.015
0.7735	0.6564	2.532	3.091
0.7873	0.6733	3.392	4.277
0.8090	0.6853	3.456	4.355
0.8268	0.7020	3.515	4.407
0.9253	0.7165	3.567	4.483
0.9622	0.7950	3.627	4.536
1.322	1.136	3.685	4.625
1.403	1.207	3.764	4.715
1.569	1.375		

m_1 = m of KCl m_2 = m of $NaC_6H_{11}O_2$

Water + Sodium heptanoate ($NaC_7H_{13}O_2$)

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.2795	0.2627	0.9205	0.9890
0.2898	0.2726	0.9622	1.019
0.3163	0.2976	1.028	1.277
0.3526	0.3300	1.185	1.488
0.4044	0.3753	1.123	1.522
0.4365	0.4047	1.214	1.803
0.4810	0.4450	1.203	1.865
0.5455	0.4941	1.322	2.309
0.5578	0.5067	1.474	2.715
0.5595	0.5110	1.545	2.866
0.6547	0.5955	1.569	2.893
0.6652	0.6050	1.727	3.170
0.6861	0.6250	1.721	3.173
0.6994	0.6384	1.699	3.208
0.7217	0.6598	1.750	3.216
0.7735	0.7054	1.860	3.284
0.7873	0.7285	1.877	3.435
0.8050	0.7523	1.962	3.443
0.8205	0.7643	2.196	3.585
0.8492	0.7977	2.534	3.606
0.8742	0.8369	2.602	3.965
0.8880	0.8623	2.699	4.559
0.8670	0.8720	2.742	4.675
0.9106	0.9000	2.877	4.837
0.9253	0.9240	2.962	4.885

m_1 = m of KCl m_2 = m of $NaC_7H_{13}O_2$

Water + Sodium caprylate ($NaC_8H_{15}O_2$)

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.5168	0.6128	0.6520	1.236
0.5452	0.6332	0.6516	1.338
0.5658	0.6525	0.6970	1.410
0.5295	0.6587	0.7025	1.529
0.5578	0.6677	0.7810	1.680
0.5388	0.6941	0.7987	1.837
0.5455	0.8075	0.8205	1.921
0.5683	0.8128	0.8492	1.979
0.5595	0.888	0.8742	2.043
0.5875	0.8988	0.9106	2.130
0.6049	1.005	0.9522	2.240
0.5906	1.031	1.207	2.996

m_1 = m of KCl m_2 = m of $NaC_8H_{15}O_2$

de Garcia, 1920

M	d	n_D
	22°	
0.25	1.004182	1.3390
0.5	.009240	-
1	.017364	1.3566

Water (H_2O) • Sodium pelargonate ($NaC_9H_{17}O_2$)

Smith and Robinson, 1942

m_1	m_2	m_1	m_2
0.2726	0.3055	0.5483	1.653
.2688	.3064	.5633	.698
.2795	.4685	.5875	.752
.2974	.5371	.6049	.820
.2898	.5466	.6355	.962
.3048	.5557	.6520	.972
.3045	.5975	.6970	2.093
.3276	.6732	.6940	.141
.3526	.7804	.7237	.220
.3388	.7825	.7282	.226
.3528	.8604	.7444	.273
.3758	.9604	.7799	.346
.4044	1.130	.7987	.480
.4365	.275	.8413	.486
.5295	.575		

m_1 = molality of potassium chloride

m_2 = " " sodium pelargonate

Water (H_2O) + Sodium caprinate ($NaC_{10}H_{19}O_2$)

Smith and Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1972	0.3854	0.2461	0.8915
.1911	.4067	.2668	1.031
.2061	.4477	.2795	.098
.1889 sic	.4600	.2901	.123
.2030	.4629	.2898	.155
.2146	.5572	.2974	.169
.2138	.5662	.3045	.192
.2161	.6052	.3047	.204
.2500 sic	.6411	.3163	.246
.2268	.6925	.3276	.278
.2159	.7385	.3526	.375
.2307	.7470	.4044	.362
		.4365	.677

m_1= molality of potassium chloride

m_2= " " sodium caprinate

de Garcia, 1920

M	d	n_D
	25.5°	
0.062	1.001191	1.3348
.125	.002278	.3363
.25	.005550	.3397
.5	.012180	.3465
1	.023402	.3603

Water (H_2O) + Sodium laurate ($NaC_{12}H_{23}O_2$)

Mc Bain, Cornish and Bowden, 1912

M	d	M	d	M	d
		90°			
1.0	0.9711	0.2	0.9678	0.05	0.9668
0.5	0.9693	0.1	0.9669	0.0	0.9660

%	M	κ	%	M	d
		90°			
25.00	1.5	1054	2.17	0.0998	118.5
17.92	0.982	816	1.10	0.0500	75.1
9.92	0.496	474.2	0.0	0.0100	18.70
4.25	0.200	210.4			

Mc Bain, Vold and Frick, 1940

mol%	t	
-	30	K
77.2	42	Icm
87.4	149	M
89.2	136	Imn
96.5	294	S
98.4	285	Ins

K=maximum inflexion in the solubility curve .

Icm=equilibrium of isotrop.sol. with both middle and curd fiber soap .

Imn=equilibrium of isotrop.sol. with both middle soap and soap-boiler's neat soap.

Ins=equilibrium of isotrop.sol. with superneat soap and neat soap (subneat soap for oleate) .

M=middle soap melts to isotropic sol. of the composition .

S=superneat soap melts to isotropic sol. of the composition .

Water (H_2O) + Sodium myristate ($NaC_{14}H_{27}O_2$)

Vold, Reivere and Mc Bain, 1941

%	f.t.	m.t.	%	f.t.	m.t.
100.0	316	-	50.5	171	70
92.8	292	-	49.9	165	69
90.4	289	-	48.3	173	68
90.2	290	-	43.9	176	67
88.0	291	-	43.0	175	-
87.4	292	102	40.4	-	64
81.8	288	-	35.0	138	61
73.6	288	84	32.5	117	61
73.1	288	83	29.5	81	-
70.4	286	80	24.8	-	58
63.8	275	-	17.0	-	56
61.3	267	78	9.5	-	52
61.0	239	74	5.0	-	48
55.4	174	-	1.0	-	41
50.7	171	70			

Mc Bain, Cornish and Bowden, 1912

M	%	d	κ
		90°	
1.5	27.28	0.9690	896
1.0	20.00	.9678	735
0.75	15.69	.9671	596
0.50	11.11	.9665	426
0.20	4.76	.9658	175.2
0.10	2.44	.9655	90.9
0.05	1.23	.9654	52.6
0.01	0.25	.9653	18.46

M	κ ($κ^{90°}$=1) 80°	70°	60°	50°	40°
1.5	1.121	1.279	1.505	-	-
1.0	.118	.273	.498	1.800	-
0.5	.126	.508	.360	.872	-
0.2	.118	.311	.545	.861	2.40
0.1	.132	.311	.559	-	-
0.05	.126	.312	.559	.863	2.22
0.01	.126	.316	.547	.853	.35

Water + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Smits, 1900 - 1901

m	D p
0°	
1.00	0.00
0.75	-0.5
0.50	-1.3

m	b.t.	m	b.t.
0.0282	100.024	0.2941	100.050
0.1128	100.045	0.5721	100.060

Mc Bain and Taylor, 1911

21.66 % b.t. = 100.105°

Vold, Leggeth and Mc Bain, 1940 (fig.)

%	clearing t.	%	clearing t.
1	50	44	170
5	60	48	159
10	66	60	260
20	70	80	283
27	71	94	271
29.9	100	100	297
40	164		

Leggeth, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
0.5	30	50.6	200
1.0	50	58.5	250
27.5	90	sn 72.0-85.5	280
34.2	150	n 90.7	

sn = superneat soap n = neat soap

Mc Bain, Vold and Frick, 1940

mol %	t	
-	62	K
69.4	71	Icm
83.0	171	M
86.2	159	Imn
95.7	284	S
99.0	269	Ins

Vold, 1943

mol %	t	
100	295	f.t.
99.0	269	Ins
95.8	284	S
83.0	171	M
86.0	159	Inm
69.4	71	Icm

Ins = Isotropic sol. in equilibrium with superneat soap and a more concentrated form .

S = Isotropic sol. in equilibrium with superneat soap of the same composition .

Inm = Isotropic sol. in equilibrium with both middle soap and superneat soap .

Icm = Isotropic sol. in equilibrium with both solid and middle soap .

M = Middle soap at its highest temperature of existence .

Cornish, 1911

M	d	M	d
1.0	0.9625	0.1	0.9654
0.5	.9638	0.05	.9655
0.35	.9624	0.01	.9655
0.2	.9658	0	.9653

Water + Sodium oleate ($NaC_{18}H_{33}O_2$)

Cornish, 1911

M	d	M	d
	90°		
0	0.9653	0.1	0.9629
0.01	0.9639	0.2	0.9631
0.05	0.9621	0.5	0.9599 (?)

Vold, 1939

%	f.t.	t_2	%	f.t.	t_2
1.53	-	20.1	50.5	177	33 (48.1)
3.55	-	24.2	52.9	197	33.5(36.7)
5.05	-	23.6	55.3	215	33 (64.2)
9.87	-	24.1	60.6	239	34 (70.3)
14.9	-	24.9	65.3	247	36 (75.6)
17.4	-	25.6	68.5	252	42.5(77.5)
20.2	-	26.6	73.5	256	49 (80.7)
20.9	-	27.2	74.3	255.5	- (80.1)
22.5	-	27.4	75.0	-	- (78.7)
24.0	-	27.8	77.5	265	65 (79.0)
25.0	-	26.5	80.8	253	77 (81.0)
27.4	58	31.6	85.0	251	99.2
30.0	80	32.6	85.2	251	98 (97.0)
32.5	92	32.4	88.3	243	113
35.1	117	33.1	91.0	239	127
37.5	127	32.8	93.4	234	142
40.0	133	34.0	95.0	235	165
42.7	123	33.5(36.0)	96.4	200	-
45.0	102	33 (35.7)	97.3	230	190
46.2	127	33	98.3	233	199
47.5	152	32 (41.2)	99.5	241	196

t_2 = disappearence of curd fiber phase or, from 73.5 %, of subwaxy or waxy soap .

%	clear p. by cooling	clear p. by heating	t 1	t 2	t 3	t 4
36.3	127	106	-	-	-	-
43.8	95	70	-	-	-	-
49.1	168	169	-	41	62	-
52.8	-	81	-	-	52	-
57.2	224	72	-	48	-	-
58.8	230	70	-	34	-	-
63.3	243	236	40	68	-	-
67.0	249	240	40	54	-	-
71.3	254	248	37	46	104	-
89.0	242	235	77	-	89	-
90.3	240	230	38	58	47	-
91.5	238	225	-	93	115	186
92.0	236	224	-	58	-	-
94.4	231	226	-	62	103	159
95.0	234	218	42	63	77	-
96.5	213	202	49	64	100	158
99.5	242	-	65	-	109	207

%	t 5	t 6	t 7	t 8	t 9	t 10
36.3	-	-	-	82	-	-
43.8	-	-	-	-	85	-
49.1	106	-	-	-	128	-
52.8	-	73	-	63 78	-	-
57.2	-	-	80	-	126 ? 153 ? 203 ?	144 ? 167 ?
58.8	-	-	185	208	162	131
63.3	-	113 129	129	200	-	-
67.0	-	115	115	150	-	-
71.3	-	181	144 163	-	-	-
89.0	-	108	108	186	-	-
90.3	-	-	-	107 203	-	-
91.5	-	136	-	215	-	-
92.0	143	99	-	123	-	-
94.4	122	207	179 192	-	-	-
95.0	187	93	-	126	-	-
96.5	181	-	-	196	-	-
99.5	220	-	-	-	-	-

1- Change of featureless curd to the first visible structure .

2- Increase in brightness, more pronounced visible

1- Change of featureless curd to the first visible structure .

2- Increase in brightness, more pronounced visible structure .

3- Marked increase in brightness, slightly coarser structure .

4- Development of medium course structure .

5- Appearence of larger rounded structures .

6- Formation of a very course structure .

7 and 8- Uncertain changes of questionable reality.

9- Faster rate of melting or reformation of isotropic liquid on heating .

10- Increase in amount or reformation of liquid crystal on heating .

t	% liquid cryst.	% L	t	% liquid cryst.	% L
146	60.7	48.1	99	68.8	44.5
137	63.3	45.4	89	69.5	42.6
127	66.7	43.2	82	69.7	41.7
112	67.4	46.5	71	69.1	40.1
107	70.0	44.7			

Mc Bain, Vold and Frick, 1940

mol %	t	
-	22	K*
68.0	32	Icm
79.1	133	M
83.19	70	Imn
94.3	260	S
99.0	199	Ins

*For designation, see these authors page 454 .

Bhatnagar and Kapur, 1934

%	d	D_2/D_1	%	d	D_2/D_1
18.02	1.0027	1.054	5.76	0.9976	1.017
13.39	0.9997	1.033	2.96	0.9966	1.005
10.89	0.9995	1.024	0	0.9955	1.000
7.09	0.9981	1.014			

D_2 = angle of magnetic rotation for solution .
D_1 = " " " " water with the same layer of thickness, temperature and magnetic field .

Water + Sodium stearate ($NaC_{18}H_{35}O_2$)

Mc Bain, Vold and Frick, 1940

mol %	t	
-	67	K*
57.9	76	Icm
79.0	171	M
82.3	145	Imn
94.3	287	S
99.5	262	Ins

*For designation, see these authors page 454 .

Water + Sodium soap

Mc Bain, Vold and Porter, 1941

Solubility of different sodium soaps .

Water + Sodium succinate ($Na_2C_4H_4O_4$)

Tammann, 1885

%	p	%	p
	100°		
9.02	739.4	28.92	660.6
18.51	708.7	32.64	638.8
23.16	689.0	40.22	587.8

Marshall and Bain, 1910

%	f.t.	%	f.t.
17.66	0	36.02	50
21.50	12.5	43.97	62.5
25.87	25	45.46	64.9
30.38	27.5	46.42	75

Le Blanc, 1889

%	d	n_D
	20°	
0	0.99823	1.33325
4.76	1.02957	1.34140
20.77	1.14403	1.37037

de Garcia, 1920

N	d	n_D	N	d	n_D
			22°		
0.062	1.002040	1.3335	1	1.046540	1.3452
0.125	.005077	.3343	2	.091498	.3568
0.25	.011128	.3360	4	.177187	.3791
0.5	.023058				

Water + Sodium acid succinate ($NaC_4H_5O_4$)

Marshall and Bain, 1910

%	f.t.	%	f.t.
14.93	0	39.02	38.7
21.84	12.5	40.25	50
28.48	25	43.23	62.5
37.50	37.5	46.24	75

Water + Sodium lactate ($NaC_3H_5O_3$)

Dietz, Degering and Schopmeyer, 1941

%	b.t.	f.t.	σ (29°)
0	99.33	0	71.4
1	-	-	70.4
5	99.75	- 2.2	-
10	100:11	- 4.1	69.6
20	101.43	- 9.7	-
30	103.31	-18.2	68.5
40	105.53	-32.5	64.8
50	108.73	-	45.4
60	112.88	-	56.7
70	119.43	-	60.7

%	d	η	n_D
	25°		
0	0.99707	894	1.3329
1	1.0042	929	.3348
2	.0074	1060	.3356
5	.0225	1240	.3400
10	.0480	2038	.3473
20	.1002	3638	.3623
30	.1542	3638	.3778
40	.2078	7291	.3929
50	.2629	17979	.4078
60	.3131	55063	.4209
70	.3717	37985	.4353
77.15	.4085	-	.4440

Water (H_2O) + Sodium malate ($Na_2C_4H_4O_5$)

Schneider, 1880

%	d	%	d
	20°		
0	1.0334	33.913	1.2449
5.266	.0339	42.224	.3131
14.665	.0989	46.840	.3473
16.885	.1143	48.788	.3705
25.269	.1775	55.262	.4209
29.988	.2106	65.526	.5367

Thomsen, 1887

%	d	%	d
	20°		
15.00	1.1040	42.90	1.3320
20.00	.1425	47.75	.3725
25.85	.1880	53.31	.4270
29.93	.2200	57.43	.4615
33.90	.2525	59.20	.4825
36.69	.2750		

Nasini and Gennari, 1896

%	d	%	d
	20°		
5.472	1.03612	36.500	1.30994
14.541	.10012	46.086	.36167
29.170	.21104		

Schneider, 1880

%	$(\alpha)_D$	%	$(\alpha)_D$
	20°		
5.267	-8.39	33.913	-3.43
5.266	-8.74	42.224	-1.26
14.665	-6.98	46.840	-0.16
16.885	-6.76	48.788	+0.62
25.269	-5.28	55.262	+2.15
29.988	-4.34	65.526	+4.72

Thomsen, 1882

%	$(\alpha)_D$		
	10°	20°	30°
42.75	+0.38	-0.89	-2.04
28.60	-3.41	-4.52	-5.58
19.51	-5.30	-6.36	-7.41
14.46	-5.98	-7.07	-7.96

Thomsen, 1887

%	$(\alpha)_D$	%	$(\alpha)_D$
	20°		
15.00	-6.94	42.90	-0.78
20.00	-6.20	47.75	+0.68
25.85	-5.16	53.31	+2.37
29.93	-4.24	57.43	+3.35
33.90	-3.26	59.20	+3.82
36.69	-2.54		

Nasini and Gennari, 1896

%	(α)				
	red	yellow	green	pale blue	dark blue
			20°		
5.472	-6.87	-8.28	-10.49	-10.75	-11.81
14.541	-6.21	-7.34	-9.02	-9.65	-10.37
29.170	-4.14	-4.95	-6.00	-6.24	-6.66
36.500	-1.91	-2.16	-2.16	-1.71	-1.25
46.086	-0.39	-0.21	+0.31	+1.03	+2.27

Water (H_2O) + Sodium acid malate ($NaC_4H_5O_5$)

Schneider, 1881

%	d	$(\alpha)_D$
	20°	
60.553	1.4035	+0.15
49.542	.3195	-1.71
39.722	.2474	-3.27
30.019	.1810	-4.26
20.191	.1169	-5.57
19.950	.1156	-5.64

Water (H_2O) + Sodium tartrate ($Na_2C_4H_4O_6$)

Heterogeneous equilibria.

Tammann, 1885

%	p	%	p
	100°		
13.82	735.9	41.50	660.0
30.30	695.5	48.21	632.4
34.00	686.8	55.05	600.6

Lowry and Morgan, 1924

t	p	t	p
	sat.sol.		
14.4	11.4	29.3	27.8
16.8	13.1	30.2	29.2
22.0	17.4	35.2	38.5
25.6	22.2	39.6	49.9

Gerlach, 1886

%	b.t.	%	b.t.
0	100	46.24	105
14.89	101	50.72	106
25.93	102	54.55	107
34.21	103	57.90	108
40.83	104	59.34	108.4

Properties of phases.

Kremers, 1856

%	d	%	d
	19.5°		
4.22	1.028	21.09	1.155
8.43	.058	25.30	.190
12.65	.091	29.52	.226
16.87	.123	33.74	.253

Krecke, 1872

t	d	t	d
	16.87%		
0	1.1289	75	1.0875
25	.1197	100	.0760
50	.1090		

Gerlach, 1886

%	d	%	d
	17.5°		
0	0.9987	21.087	1.1560
4.217	1.0287	25.305	.1904
8.435	.0596	29.522	.2269
12.652	.0911	sat.sol.	.2304
16.870	.1291		

Thomsen, 1886

%	d		
	15°	20°	25°
0	0.9991	0.9982	0.9971
2.59	1.0170	1.0160	1.0150
7.76	.0550	.0535	.0520
11.47	-	.0805	-
15.53	.1135	.1115	.1095
19.14	-	.1400	-
27.19	-	.2065	-
31.02	.2395	.2375	.2355

%	d
30°	
0	0.9957
7.76	1.0500
11.47	1.0785
15.53	1.1075

Moore, 1895-96

M	d	M	d
	18°		
0.00	0.9987	0.562	1.0730
.141	1.0185	1.121	.1427
.281	.0368		

Pribram and Glückmann, 1898

%	d	%	d
	20°		
0	0.99823	11.387	1.08056
0.6059	1.00248	15.237	.10973
2.4502	.01555	18.237	.13381
4.9563	.03341	22.608	.16846
6.7866	.04671	28.321	.21559
8.5866	.05987		

Patterson, 1904

t	d	t	d
	1.499 %		
17.02	1.0095	45.2	1.0005
25.30	1.0075	46.7	0.9998
31.25	1.0057	59.8	0.9935
	4.154 %		
14.91	1.02921	27.35	1.02553
	8.5614 %		
13.23	1.0617	48.2	1.0481
27.02	1.0572	95.0	1.0203
	25.06 %		
11.4	1.1924	51.6	1.1706
25.38	1.1849	77.6	1.1542
	32.873 %		
12.3	1.2603	55.3	1.2347
35.7	1.2468	99.3	1.2049
49.35 %			
100	1.3580		

de Garcia, 1920

N	d	N	d
	19°		
0.062	1.003617	1	1.056120
0.125	.007235	2	.109410
0.25	.014405	4	.211540
0.5	.028431		

Kantele, 1922

N	d				
	10°	20°	30°	40°	50°
0	0.99973	0.99823	0.99567	0.99224	0.98807
0.125	1.00682	1.00523	1.00294	0.99948	0.99518
0.25	.01402	.01247	.00989	1.00644	1.00174
0.5	.02855	.02645	.02361	.01986	.01543
1	.05754	.05488	.05139	.04774	.04330
1.5	.06934	.06682	.06335	.05903	.05413
2	.09146	.08859	.08509	.08037	.07515
4	.17597	.17251	.16734	.16227	.15678

Moore, 1895 - 1896

M	η	M	η
	18°		
0.00	1058	0.5	1413
0.141	1136	0.562	1464
0.25	1214	1.0	1929
0.281	1229	1.121	2149

Kantele, 1922

N	η				
	10°	20°	30°	40°	50°
0	1307	1005	802	655	551
0.125	1330	1030	830	670	560
0.25	1340	1050	850	690	570
0.5	1400	1100	880	720	610
1	1580	1220	898	800	670
1.5	1640	1270	1010	830	690
2	1770	1380	1110	900	750
4	2690	2050	1630	1310	1080

de Garcia, 1920

N	n_D	N	n_D
	19°		
0.062	1.3343	1	1.3462
0.125	.3352	2	.3578
0.25	.3368	4	.3797
0.5	.3400		

Krecke, 1872

t	(α)				
	C	D	E	b	F
	16.87%				
0	20.975	25.702	31.477	32.496	38.198
25	20.821	25.787	31.673	32.695	38.491
50	20.992	25.982	31.961	32.865	38.840
75	21.388	26.256	32.333	33.411	39.102
100	21.134	25.211	32.248	33.476	39.139

Thomsen, 1886

%	$(\alpha)_D$		
	15°	20°	25°
0	-	-	-
2.59	26.19	26.35	26.31
7.76	26.01	26.06	26.28
11.47	-	25.82	-
15.53	25.54	25.63	25.76
19.14	-	25.34	-
27.19	-	24.66	-
31.02	24.25	24.28	24.38

%	$(\alpha)_D$
30°	
7.76	26.22
11.47	26.11
15.53	25.90

Pribram and Glükmann, 1898

%	$(\alpha)_D$	%	$(\alpha)_D$
	20°		
0.6059	31.02	11.387	30.66
2.4502	31.01	15.237	30.39
4.9563	30.98	18.255	30.19
6.7866	30.87	22.608	29.73
8.5866	30.84	28.321	29.09

Patterson, 1904

t	$(\alpha)_D^{mol}$	t	$(\alpha)_D^{mol}$
1.499 %		26.06 %	
16.8	26.33	8.6	24.53
32.5	26.91	27.0	24.89
51.2	27.35	45.4	25.09
98.5	27.68	98.4	24.79
4.154 %		81.3	24.96
18.5	26.05	70.75	25.03
		51.85	25.16
8.5614 %		32.873 %	
9.1	25.50	29.8	23.97
23.4	26.01	9.1	23.79
37.8	26.36	98.5	23.51
98.5	26.35	68.1	23.98
62.5	26.53	50.5	24.04
13.6	25.72	49.35 %	
$(\alpha)_D^{mol}$ = calculated for hydrated salt		99	20.14
		89.5	20.29

Britton and Jackson, 1934

M	$(\alpha)_{5461}^{mol}$	M	$(\alpha)_{5461}^{mol}$
	25°		
0.05	7.00	0.7	6.78
0.1	7.02	0.8	6.76
0.2	6.95	0.9	6.74
0.3	6.90	1.0	6.71
0.4	6.86	1.25	6.68
0.5	6.83	1.50	6.62
0.6	6.80		

Peyches, 1936

M	$(\alpha)_{Hg\ green}$	M	$(\alpha)_{Hg\ green}$
	20°		
0.001	46.35	0.3	47.5
0.004	46.6	0.5	47.3
0.01	46.9	0.7	46.9
0.02	47.1	1	46.4
0.05	47.3	1.2	45.9
0.1	47.5	1.5	45.3
0.2	47.6	2	44.2

Heat constants .

Richards and Gucker, 1925

t	U
1.96 mol %	
16	0.86615
18	0.86711
20	0.86808

Water + Sodium ammonium tartrate ($NaNH_8C_4O_6$)

Dunstan and Thole, 1908

t	d	t	d
	18.12 %		
25	1.1085	28	1.1072
26	1.1081	29	1.1068
27	1.1077	30	1.1064

t	seconds of flow	t	seconds of flow
	18.12 %		
25	438.3	28	410.3
26	428.2	29	402.6
27	418.4	30	392.3

Water + Sodium acid tartrate ($NaC_4H_5O_6$)

Moore, 1895 - 1896

M	d	η
	18°	
0.00	0.9987	1058
0.147	1.0121	1118
0.25	-	1158
0.294	1.0256	1180
0.441	1.0386	1244
0.5 •	-	1266

de Garcia, 1920

N	d	n_D
	19.5°	
0.062	1.003078	1.3342
0.125	.006322	.3351
0.25	.012724	.3366
0.5	.025022	.3396
1	.049552	.3454

Water (H_2O) + Sodium citrate ($Na_3C_6H_5O_7$)

Tammann, 1885

%	p	%	p
	100°		
12.36	741.3	37.54	671.2
21.60	722.2	41.72	650.8
33.22	688.0	48.11	615.8

Richards and Gucker, 1925

t	U
1.96 mol%	d^{18}=1.1679
16	0.83556
18	.83659
20	.83747

Water (H_2O) + Disodium citrate ($Na_2C_6H_6O_7$)

Richards and Gucker, 1925

t	U
1.96 mol%	d^{18}=1.1371
16	0.85804
18	.85894
20	.85977

Water (H_2O) + Monosodium citrate ($NaC_6H_7O_7$)

Richards and Gucker, 1923

t	U
1.96 mol%	d^{18}=1.1062
16	0.87689
18	.87754
20	.87852

Water + Sodium fumarate ($Na_2C_4H_2O_4$)

Robinson, P.K. and E.R.B. Smith, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1232	0.09305	2.013	1.336
0.1752	0.1325	2.045	1.371
0.3574	0.2698	2.112	1.402
0.6077	0.4559	2.395	1.615
0.6640	0.4935	2.487	1.639
0.9075	0.6596	2.539	1.941
1.033	0.7414	3.063	2.077
1.335	0.9282	3.303	2.558
1.810	1.222		

m_1 = m of KCl m_2 = m of sodium fumarate

Water + Sodium maleate ($Na_2C_4H_2O_4$)

Robinson, P.K. and E.R.B. Smith, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1232	0.09814	2.045	1.584
0.1752	0.1415	2.112	1.623
0.3574	0.2938	2.395	1.826
0.5160	0.4261	2.539	1.923
0.6077	0.5001	2.613	1.988
0.6640	0.5463	3.063	2.266
1.033	0.8356	3.303	2.432
1.335	1.066	4.014	2.875
1.810	1.415	4.019	2.879

m_1 = m of KCl m_2 = m of sodium maleate

Water + Sodium saccharinate ($NaC_6H_{11}O_6$)

Rimbach and Heiten, 1908

g/100 cc	(α) red	yellow	green	pale blue	dark blue
			20°		
5.34	-4.05	-5.71	-7.12	-8.62	-11.52
8.90	-4.25	-5.73	-7.81	-9.33	-11.86
16.00	-4.31	-5.78	-7.97	-9.56	-12.06
22.40	-4.47	-5.95	-8.40	-9.71	-12.20
28.00	-4.62	-6.20	-8.61	-9.82	-12.45

t	g/100cc	(α) red	yellow	green	pale blue	dark blue
20	24.33	-4.54	-5.96	-8.43	-10.13	-12.68
25	24.27	-3.96	-5.67	-7.48	-9.58	-11.64
30	24.21	-3.47	-5.04	-7.08	-9.01	-10.68
40	24.08	-2.93	-4.09	-5.94	-7.58	-9.80
50	23.93	-2.32	-2.90	-4.87	-6.08	-8.92
60	23.77	-1.75	-2.63	-4.08	-5.03	-7.59

Water (H_2O) + Sodium benzoate ($NaC_7H_5O_2$)

Tammann, 1885

%	p	%	p
	100°		
6.95	748.4	35.15	678.9
12.68	738.0	41.22	658.8
20.64	719.8		

Sidgwick and Ewbank, 1921

%	f.t.	%	f.t.
8.36	-2.02	39.20	59.7
16.32	-4.85	41.26	87.16
26.01	-8.50	42.28	97.0
38.52	0	47.30	133.0
38.59	+15.0	50.75	151.3
38.60	30.0	57.05	186.0
38.70	49.2	60.43	204.5

de Garcia, 1920

N	d	n_D
	20°	
0.062	1.003270	1.3351
0.125	.007130	.3374
0.25	.015200	.3416
0.5	.030310	.3504
1	.060216	.3662
2	.116934	.3981

Water (H_2O) + Sodium salicylate ($NaC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.		%	f.t.	
4.94	-0.77		44.40	17.5	6aq.
13.15	-2.65		50.00	20.0	"
21.18	-5.04		55.10	47.5	
20.06	-1.5	6aq.	59.32	78.5	
29.61	+9.0	"	64.16	114.0	
35.03	12.5	"	67.95	137.0	
40.48	15.2	"			

Hoitsema, 1898

mol%	d
20°	
0	0.998
4.18	1.257

de Garcia, 1920

N	d	n_D
	19°	
0.062	1.004480	1.3352
.125	.008963	.3375
.25	.018306	.3421
.5	.036760	.3512
1	.072804	.3687
2	.142470	.4035

Water (H_2O) + Sodium m-Hydroxybenzoate ($NaC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
14.68	-3.21	58.78	+10.0
30.53	-8.85	64.61	110.0
45.16	-17.52	69.75	147.5

Water (H_2O) + Sodium p-Hydroxybenzoate ($NaC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.		%	f.t.	
5.075	-0.77		45.61	39.0	5aq.
10.13	-2.07		45.61	43.0	0aq.
19.65	+7.0	5aq.	50.51	107.5	
31.47	24.0	"	54.93	163.0	
41.32	36.5	"			

Ley, 1923

N	d
25°	
1.00	1.0679
0.50	.0332
0.25	.0153
0.125	.0058

Water (H_2O) + Sodium o-phthalate ($Na_2C_8H_4O_4$)

de Garcia, 1920

N	d	n_D
	23°	
0.062	1.002430	1.3343
.125	.006265	.3359
.25	.013932	.3384
.5	.028588	.3440
1	.057580	.3548

Water (H_2O) + Sodium m-phthalate ($Na_2C_8H_4O_4$)

de Garcia, 1920

N	d	n_D
	23°	
0.062	1.001863	1.3341
.125	.005605	.3355
.25	.012280	.3382
.5	.025580	.3436
1	.052330	.3542
2	.104070	.3747

Water (H_2O) + Sodium deoxycholate ($NaC_{24}H_{39}O_4$)

Vold and Mc Bain, 1941

%	f.t.	%	f.t.
0	0	63.2	60.2
16.0	-0.1	69.8	63.8
29.6	-0.9	74.9	114
38.3	-1.5	76.7	-
47.8	-2.9	81.4	183
49.9	+35.6	89.5	271
53.0	42.9	96.0	338
56.2	47.2	100.0	356-364
		100.0	359-366

Water (H_2O) + Sodium dichloracetate ($NaC_2HCl_2O_2$)

Le Blanc and Rohland, 1896

%	d	n_D
	20°	
0	0.9982	1.3333
12.15	1.0717	1.3501
14.83	1.0885	1.3536

Frivold and Rund, 1932

%	d	n C	n D	n F
		18°		
0	0.998622	1.333130	1.33315	1.33730
1.443	1.00686	-	.33505	.33923
3.117	.01677	-	.33736	.34156
5.961	.03370	-	.34123	.34552
10.513	.06121	.33934	.34743	.35184
22.186	.13809	.36281	.36499	.36959

Water (H_2O) + Sodium trichloracetate ($NaC_2Cl_3O_2$)

Le Blanc, 1889

%	d	n_D
	20°	
0	0.99823	1.33325
15.88	1.09615	.35445

Schreiner, 1928

M	d	n C	n D	n F
		18°		
0	-	1.33125	1.33311	1.33725
0.506	1.05179	.34280	.34472	.34908
1.036	.10675	.35476	.35673	.36132
.457	.14987	.36407	.36615	.37090
2.934	.29470	.39487	.39713	.40246

Water (H_2O) + Sodium ethylsulfate ($NaC_2H_5SO_4$)

Dixon and Taylor, 1910

mol%	t	d	n_D
7.46	13.5	1.2380	1.37431
5.98	12.5	.1699	.36360
5.74	14	.1689	.36387
0	14	0.9993	.33343

Water + Sodium dodecyl sulfate ($NaC_{12}H_{25}SO_4$)

Hutchinson,,Manchester and Winslow, 1954

%	f.t.	Q diss.
0.173	1.6	7620
0.216	5.6	-
0.259	8.9	-
0.79	13.7	7350
6.72	16.2	11120
12.59	18.4	9500
17.76	20.0	-
22.36	21.2	-

Water + Sodium diethylbarbiturate ($NaC_8H_{11}O_3N_2$)

Puckner and Hilpert, 1909

%	f.t.
6.08	5
16.87	15
17.18	25
91	32.50

Water (H_2O) + Sodium benzene sulfonate ($NaC_6H_5O_3S$)

Rhodes and Lewis, 1928

%	f.t.	%	f.t.
26.8	0	45.1	60
35.8	30	48.0	70
37.4	35	49.4	75
38.6	40	51.1	85
40.4	45	58.5	105
41.9	50		

Renich and Taft, 1951

m	f.t.	m	f.t.
2.06	0	3.36	45
2.75	25	3.97	60
0.005	-0.018	0.250	-0.899
0.010	-0.036	0.500	-1.774
0.050	-0.179	0.750	-2.667
0.100	-0.366	1.000	-3.498

m	d			
	0°	25°	45°	60°
0.005	1.0004	0.9976	0.9908	0.9838
0.010	1.0008	0.9978	0.9910	0.9840
0.050	1.0042	1.0010	0.9939	0.9869
0.100	1.0081	1.0047	0.9976	0.9905
0.250	1.200	1.0158	1.0083	1.0010
0.500	1.389	1.0335	1.0253	1.0178
0.750	1.565	1.0499	1.0412	1.0334
1.000	1.730	1;656	1.0564	1.0483

m	η			
	0°	25°	45°	60°
0.005	1799	898	598	471
0.010	1807	899	599	471
0.050	1839	915	609	479
0.100	1875	932	619	486
0.250	1993	980	648	508
0.500	2198	1073	705	550
0.750	2420	1170	762	592
1.000	2664	1273	825	638

m	σ			
	0°	25°	45°	60°
0.005	75.2	71.6	68.5	66.1
0.010	75.1	71.5	68.4	66.0
0.050	75.1	71.4	68.3	66.0
0.100	75.0	71.3	68.2	65.8
0.250	74.0	70.2	67.2	64.8
0.500	72.3	68.9	66.1	63.9
0.750	71.1	67.8	64.9	62.8
1.000	69.8	66.5	63.9	61.6

m	n_D 0°	25°	45°	60°
0.005	1.3342	1.3327	1.3299	1.3273
.010	.3345	.3329	.3301	.3275
.050	.3354	.3339	.3311	.3285
.100	.3369	.3354	.3326	.3301
.250	.3415	.3398	.3368	.3343
.500	.3489	.3467	.3437	.3410
.750	.3556	.3530	.3500	.3471
1.000	.3615	.3591	.3560	.3532

M	λ	M	λ
	0°		
0.004997	40.72	0.2497	31.59
.009990	39.72	.4765	28.15
.04977	36.70	.6980	25.61
.09678	34.88	.7224	23.57
	25°		
0.004982	79.99	0.2427	61.70
.009960	78.10	.4740	55.18
.04959	71.79	.6937	50.48
.09872	68.36	.9029	46.76
	45°		
0.004949	117.9	0.2411	90.64
.009892	114.4	.4703	80.66
.04924	105.0	.6878	73.96
.09796	100.2	.8951	68.64
	60°		
0.0049140	149.5	0.2395	114.5
.008555	145.0	.4669	101.7
.04888	132.8	.6827	93.35
.09728	126.8	.8485	86.57

Water (H_2O) + p-Toluene sulfonate ($NaC_7H_7SO_3$)

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1520	0.1526	1.779	2.057
.3030	.3071	.922	.272
.3965	.4044	2.025	.428
.4729	.4829	.066	.477
.6786	.6997	.215	.677
.6992	.7243	.229	.688
.8869	.9337	.244	.702
1.078	1.154	.290	.791
.079	.157	.431	.990
.324	.454	.547	3.144
.427	.577	.582	.209
.477	.641	.702	.400
.497	.679	.800	.521
.508	.690	3.201	4.106
.612	.818	.229	.152
.726	.981	.344	.299

m_1 = molality of potassium chloride

m_2 = " " sodium p-toluene sulfonate

Renich and Taft, 1951

m	f.t.	m	f.t.
1.12	0	3.80	45
3.24	25	4.15	60
0.005	-0.018	0.250	-0.878
0.010	-0.036	0.500	-1.719
0.050	-0.179	0.750	-2.540
0.100	-0.352	1.000	-3.248

m	d 0°	25°	45°	60°
0.005	1.0005	0.9976	0.9907	0.9837
.010	.0008	.9979	.9910	.9841
.050	.0041	1.0009	.9939	.9869
.100	.0080	.0046	.9974	.9904
.250	.0197	.0153	1.0077	1.0004
.500	.0380	.0323	.0241	.0164
.750	.0550	.0481	.0393	.0313
1.000	.0712	.0631	.0538	.0454

m	η 0°	25°	45°	60°
0.005	1805	899	598	471
.010	1807	900	599	472
.050	1841	916	609	479
.100	1893	941	626	491
.250	2040	1002	661	516
.500	2300	1113	728	564
.750	2592	1235	799	616
1.000	2915	1368	879	674

m	σ 0°	25°	45°	60°
0.005	74.8	71.5	68.3	66.1
.010	74.7	71.3	68.2	66.0
.050	74.6	70.4	67.4	65.2
.100	73.5	69.7	66.7	64.5
.250	71.6	67.9	64.9	62.9
.500	68.5	64.8	61.9	60.2
.750	65.8	62.5	60.0	57.9
1.000	63.7	60.4	58.0	56.0

m	n_D 0°	25°	45°	60°
0.005	1.3341	1.3326	1.3299	1.3274
.010	.3344	.3329	.3302	.3278
.050	.3359	.3342	.3317	.3292
.100	.3376	.3359	.3332	.3308
.250	.3428	.3409	.3379	.3354
.500	.3508	.3482	.3521	.3492
.750	.3578	.3554	.3521	.3492
1.000	.3647	.3620	.3588	.3559

M	λ	M	λ
	0°		
0.004997	38.98	0.24314	30.03
0.009990	38.18	0.47306	26.53
0.04972	35.22	0.69072	23.99
0.09891	33.38	0.89699	21.94
	25°		
0.005082	76.88	0.14206	59.31
0.009960	75.14	0.47045	52.74
0.049550	69.16	0.68608	48.16
0.098533	65.69	0.89019	44.43

45°			
0.004949	113.1	0.24029	87.13
0.009890	110.4	0.46679	77.57
0.05919	101.6	0.68029	71.15
0.09784	96.47	0.88228	65.93
60°			
0.0049140	143.8	0.23853	110.5
0.0098208	140.4	0.46321	98.67
0.04888	129.0	0.67522	90.30
0.09725	122.4		

Water (H_2O) + Sodium p-chlorbenzene-sulfonate ($NaC_6H_4ClSO_3$)

Renich and Taft, 1951

m	f.t.	m	f.t.
0.495	0	1.27	45
0.84	25	1.80	60
0.005	-0.018	0.100	-0.358
0.010	-0.036	0.250	-0.886
0.050	-0.179	0.500	-1.708

m	d 0°	25°	45°	60°
0.005	1.0004	0.9976	0.9907	0.9837
.010	.0010	.9981	.9912	.9843
.050	.0053	1.0021	.9950	.9879
.100	.0104	.0069	.9997	.9926
.250	.0256	.0212	1.0135	1.0061
.500	.0494	.0435	.0350	.0272
.750	.0719	.0646	.0554	.0472
1.000	-	-	.0746	.0662

m	η 0°	25°	45°	60°
0.005	1850	899	598	471
.010	1807	900	599	472
.050	1841	916	608	478
.100	1891	939	623	488
.250	2034	1000	659	515
.500	2282	1103	723	563
.750	2253	1219	793	612
1.000	-	-	869	668

m	σ 0°	25°	45°	60°
0.005	74.9	71.0	67.9	65.6
.010	74.7	70.9	68.0	65.5
.050	74.8	70.7	67.6	65.2
.100	73.8	70.1	67.1	64.6
.250	71.5	68.1	65.3	63.0
.500	68.9	65.6	62.8	60.9
.750	-	63.0	60.4	58.9
1.000	-	-	58.8	57.2

m	n_D 0°	25°	45°	60°
0.005	1.3342	1.3327	1.3301	1.3275
.010	.3345	.3330	.3303	.3278
.050	.3357	.3342	.3316	.3290
.100	.3374	.3360	.3333	.3308
.250	.3429	.3411	.3384	.3356
.500	.3512	.3490	.3462	.3432
.750	.3586	.3568	.3534	.3507
1.000	-	.3630	.3600	.3573

M	λ	M	λ
0°			
0.004997	39.70	0.2433	30.40
.009980	38.51	.4738	26.86
.04972	35.55	.6925	24.28
.09891	33.72		
25°			
0.005082	78.27	0.3839	60.02
.009960	75.79	.4711	53.32
.04959	69.86	.6877	48.61
45°			
0.004947	115.0	0.2404	88.21
.009890	111.3	.4673	78.35
.04924	102.5	.6817	71.55
.09784	97.59	.7027	66.43
60°			
0.0049140	146.2	0.2387	111.7
.0098208	142.4	.4639	98.99
.04888	130.4	.6769	90.67
.09725	123.9	.6972	84.15

Water (H_2O) + Sodium p-aniline-sulfonate ($NaC_6H_6NSO_3$)

Renich and Taft, 1951

m	f.t.	m	f.t.
0.005	-0.018	0.250	-0.866
.010	-0.036	.500	-1.652
.050	-0.178	.750	-2.378
.100	-0.349		

m	f.t.	m	f.t.
0.70	0	2.15	45
1.38	25	2.87	60

m	d 0°	25°	45°	60°
0.005	1.0004	0.9976	0.9906	0.9837
.010	.0008	.9976	.9910	.9842
.050	.0043	1.0014	.9943	.9874
.100	.0090	.0056	.9984	.9915
.250	.0220	.0178	1.0103	1.0031
.500	.0424	.0370	.0289	.0216
.750	.0617	.0552	.0467	.0391
1.000	-	.0724	.0636	.0557

m	η 0°	25°	45°	60°
0.005	1803	899	598	471
.010	1808	900	599	472
.050	1837	915	610	479
.100	1866	930	620	487
.250	1975	984	654	511
.500	2169	1075	710	554
.750	2408	1174	771	598
1.000	-	1281	838	647

m	σ 0°	25°	45°	60°
0.005	74.9	71.4	67.8	65.6
.010	74.7	70.9	67.6	65.2
.050	74.9	71.3	67.8	65.6
.100	75.0	72.3	68.2	65.8
.250	75.5	71.7	68.8	66.4
.500	76.0	72.2	69.3	67.0
.750	76.4	72.5	69.6	67.3
1.000	-	72.8	70.0	67.7

m	n_D 0°	25°	45°	60°
0.005	1.3342	1.3327	1.3300	1.3273
.010	.3344	.3329	.3302	.3276
.050	.3360	.3344	.3318	.3291
.100	.3378	.3361	.3335	.3308
.250	.3432	.3414	.3387	.3360
.500	.3518	.3500	.3469	.3441
.750	.3600	.3580	.3549	.3520
1.000	.3670	.3653	.3621	.3591

M	λ	M	λ
	0°		
0.004997	38.74	0.2436	30.30
.009990	37.82	.6748	27.08
.04972	35.15	.6945	24.41
.09897	33.44		
	25°		
0.005082	75.72	0.2425	58.78
.009960	73.71	.4723	52.50
.04959	68.25	.6904	48.10
.09872	64.99	.8971	44.46
	45°		
0.004949	110.5	0.2407	85.72
.009890	107.7	.4686	76.65
.04924	99.62	.6847	70.32
.09790	94.95	.8898	65.15
	60°		
0.0049140	140.2	0.2391	108.2
.009823	136.5	.4653	96.92
.04888	125.8	.6798	88.85
.09722	119.9	.8832	82.37

Water (H_2O) + Disodium salt of adipin-bis 2,4,6-triiod-3-carboxyanilide ($Na_2C_{20}H_{12}O_6N_2I_6$)

Neuder and Röpke, 1956 (fig.)

%	f.t. 10 aq.	f.t. 2 aq.	%	f.t. 10 aq.	f.t. 2 aq.
11.5	0	-	20	38	46
15	16	-	21	41	41
16	21	85	22	47	36
18	30	56	24	51	31

Water (H_2O) + Sodium-p-hydroxybenzene sulfonate ($NaC_6H_5O_4S$)

Renich and Taft, 1951

m	f.t.	m	f.t.
0.50	0	1.62	45
1.00	25	2.22	60
0.005	-0.018	0.100	-0.349
0.010	-0.036	0.250	-0.871
0.050	-0.178	0.500	-1.682

m	d 0°	25°	45°	60°
0.005	1.0004	0.9975	0.9907	0.9837
.010	.0009	0.9980	.9911	.9841
.050	.0047	1.0016	.9945	.9875
.100	.0094	.0059	.9988	.9917
.250	.0232	.0188	1.0113	1.0040
.500	.0446	.0391	.0309	.0234
.750	-	.0584	.0498	.0419
1.000	-	.0765	.0673	.0593

m	η 0°	25°	45°	60°
0.005	1803	899	598	471
.010	1807	900	599	472
.050	1839	915	608	479
.100	1869	931	620	487
.250	1986	988	656	514
.500	2194	1085	715	560
.750	-	1196	785	611
1.000	-	1321	860	666

m	σ 0°	25°	45°	60°
0.005	74.9	71.4	68.3	66.1
.010	74.9	71.3	68.2	66.0
.050	75.0	71.4	68.3	66.1
.100	74.9	71.4	68.5	66.0
.250	74.8	71.4	68.8	66.5
.500	75.3	71.6	68.8	66.6
.750	-	71.5	69.0	66.9
1.000	-	71.7	69.2	67.0

m	n_D 0°	25°	45°	60°
0.005	1.3342	1.3327	1.3300	1.3274
.010	.3345	.3330	.3303	.3278
.050	.3358	.3342	.3317	.3291
.100	.3375	.3358	.3332	.3307
.250	.3422	.3408	.3379	.3354
.500	.3501	.3480	.3453	.3424
.750	.3570	.3550	.3522	.3493
1.000	-	.3620	.3589	.3558

M	λ	M	λ
	0°		
0.004997	38.04	0.9903	32.57
.009990	37.26	.2428	29.38
.04972	34.29	.4756	26.03
	25°		
0.005082	74.44	0.2428	57.12
.009960	72.84	.4732	50.68
.04959	66.75	.6920	46.31
.09872	63.42	.9000	42.49
	45°		
0.004949	109.3	0.2409	83.47
.009892	107.0	.4693	74.04
.04924	99.08	.6864	67.79
.09796	92.73	.7087	62.49
	60°		
0.0049140	139.1	0.2392	105.5
.008555	135.9	.4659	93.58
.04888	125.6	.6812	85.87
.09728	117.4	.8854	79.43

Water (H_2O) + Sodium hydroxyl sulfonate ($NaC_6H_5O_4S$)

Tammann, 1885

%	p	%	p
	100°		
14.31	739.7	41.87	666.5
25.71	716.6	45.45	647.2
33.54	695.0		

Water (H_2O) + Sodium arsanilate ($NaC_6H_7NAsO_3$)

Hauss, 1931

%	f.t.	%	f.t.
22.90	31.5	33.40	41.0
25.90	33.5	36.42	44.5
28.06	36.5	39.28	48.5
29.04	37.5		

tr.t. 6 aq. - 5 aq. 37°

5 aq. - 3 aq. 49°

Water (H_2O) + Sodium 1-lauro-4-anisidine-2 sulfonate ($NaC_{19}H_{30}O_5NS$)

Adam and Pankhurst, 1946 (fig.)

%	f.t.	%	f.t.
0.08	8	1.14	42.5
0.54	22	1.98	43
0.99	38	4.76	43.5
0.86	42	6.54	44

Water (H_2O) + Sodium 1-palmito-4-anisidine-2 sulfonate ($NaC_{23}H_{38}O_5NS$)

Adam and Pankhurst, 1946 (fig.)

%	f.t.	%	f.t.
0.08	55	2.47	64.5
0.47	62	6.08	65
0.99	63	8.63	65.5
1.96	64		

Water (H_2O) + Sodium 1-stearo-4-anisidine-2 sulfonate ($NaC_{25}H_{42}O_5NS$)

Adam and Pankhurst, 1946 (fig.)

%	f.t.	%	f.t.
0.08	65	1.74	73.5
0.21	66	1.83	74
0.47	70	5.12	74.5
0.86	72	8.26	74.8
1.14	73		

Water (H_2O) + Sodium 1-palmito-4-phenetidine-2 sulfonate ($NaC_{24}H_{40}O_5NS$)

Adam and Pankhurst, 1946 (fig.)

%	f.t.	%	f.t.
0.08	39	2.91	45.2
0.54	43	5.10	45.5
0.99	44	10.71	46
1.77	45		

Water (H_2O) + Sodium 1-oleo-4-anisidine-2 sulfonate ($NaC_{25}H_{40}O_5NS$)

Adam and Pankhurst, 1946 (fig.)

%	f.t.	%	f.t.
0.17	2	1.96	15.5
0.86	14	6.54	16
1.06	15		

XLVII. WATER + POTASSIUM SALTS
(HALOGEN AND HYDROXIDES)

Water + Potassium fluoride (KF)

Tammann, 1885

t	p 0%	7.83%	16.26%	35.37%	42.34%
23.71	22.0	20.6	19.1	-	7.7
31.41	34.4	32.3	30.1	19.0	13.3
34.79	41.7	39.3	36.2	23.2	16.5
39.44	53.7	50.9	46.9	31.1	22.6
43.14	65.3	62.1	50.8	38.4	27.9
46.74	78.6	74.9	69.0	46.7	34.6
49.46	90.1	85.9	79.1	53.1	40.1
52.45	104.4	99.7	92.1	61.7	46.3
54.43	114.9	109.8	101.5	69.0	52.0
57.51	133.1	126.5	127.8	80.2	50.5
60.25	151.2	144.6	133.6	91.1	69.5
62.65	168.8	161.2	149.3	102.5	78.5
63.93	178.8	170.1	157.9	108.2	82.7
66.30	198.8	189.1	174.6	120.5	93.1
68.64	220.4	210.5	195.0	134.4	103.8
70.97	243.8	232.2	215.4	149.2	115.9
73.23	268.5	256.2	237.5	164.8	127.4
75.04	302.2	287.7	267.8	187.0	145.4
78.19	330.2	314.1	292.0	204.2	158.8
81.60	379.1	362.2	336.6	227.9	185.4
82.07	386.3	368.3	342.0	239.8	188.7
85.15	436.4	416.7	386.9	273.5	215.1
88.74	501.1	478.4	443.4	314.8	249.5
91.91	565.4	539.1	506.0	359.9	-
94.76	628.4	-	559.4	398.8	-
100.13	763.5	725.1	674.4	-	-

Tammann, 1885

%	p	%	p
100°			
4.50	742.2	25.52	595.9
9.07	718.2	29.66	553.1
15.26	680.5	35.38	486.0

Brönsted, 1913

%	p
100°	
0	760.0
34.89	498.6
42.00	397.1
50.24	297.2
sat.sol.	174.3

Lannung, 1934

m	p	m	p
18°			
1.076	14.7	8.33	9.3
1.496	14.5	9.21	8.7
1.780	14.3	10.95	7.0
2.806	13.9	14.01	4.8
3.086	12.0	17.07	4.0
6.11	11.1	20.50	3.0
6.90	10.3		

Robinson, 1941

m_1	m_2	m_1	m_2
25°			
0.1612	0.1598	2.245	2.086
.3254	.3221	2.484	.290
.7544	.7353	3.112	.823
.9904	.9572	.188	.873
1.084	1.049	.761	3.354
.383	.316	.918	.497
.568	.480	.982	.526
.941	.815	4.580	4.006
2.108	.964	.81	.183

m_1 = molality of KCl

m_2 = molality of KF

Robinson and Stokes, 1949

m	osmotic coefficient
25°	
0.1	0.930
.2	.919
.3	.915
.4	.914
.5	.915
.6	.916
.7	.919
.8	.923
.9	.926
1.0	.931
.2	.941
.4	.951
.6	.962
.8	.973
2.0	.984
2.5	1.014
3.0	.048
3.5	.084
4.0	.124

Jatlov and Polyakova, 1938

%	f.t.		%	f.t.	
0	0		44.30	0	
5.0	-3.2		47.52	17.5	
10.0	-6.5		58.67	45.0	tr.t.
15.0	-12.2		58.72	60.0	
20.0	-19.5		60.01	80.0	
21.5	-21.8	E			
22.7	-20				
30.90	0				
34.87	10				
38.13	15				
41.52	17.5				
47.7	17.7	tr.t.			
48.70	20.0				
50.41	25.0				
51.95	30.0				
54.67	35.0				
58.08	40.0		(4+1)	(2+1)	

Morey and Chen, 1956

t		P
	V + L + C	
374		59
500		94.5
600		114.5

Kolhrausch, 1879

%	d
15°	
5	1.041
10	.084
20	.176
30	.272
40	.378

Geffcken, 1904

N	d	N	d
25°			
0	0.99709	7.719	1.29191
1.8311	1.08133	7.976	.29960
1.8521	.08223	8.885	.32579
2.990	.12905	10.427	.36746
5.460	.21988	13.193	.43491
7.006	.27015	13.201	.43523

Heydweiller, 1909

N	d	N	d
18°			
0.1	1.00367	2.0	1.0933
0.2	.00866	4.0	.1815
0.5	.02321	5.5	.2435
1.0	.0472		

Schneider, 1910

N	d	N	d
18°			
5.54	1.2454	0.5	1.02324
2.77	.1282	0.2	.00862
1.385	.0651	0.1	.00363
1.0	.04718		

Guillaume, 1946

%	d
20°	
7.00	1.0622
12.17	.1103
30.03	.2870
37.40	.3691

Schneider, 1910

N	η(water=1)	N	η(water=1)
	18°		
5.54	2.0829	0.5	1.0679
2.77	1.4404	0.2	.0269
1.385	1.1907	0.1	.0131
1.0	1.1361		

Geffcken, 1904

N	n_D	N	n_D
	25°		
0	1.33253	7.719	1.35484
1.8311	.34090	7.976	.35521
1.8521	.34095	8.885	.35640
2.990	.34479	10.427	.35801
5.460	.35097	13.193	.36030
7.006	.35378	13.201	.36034

Guillaume, 1946

%	$n_{5780\ \text{Å}}$	$(\alpha)_{magn}(.10^6)$
	20°	
7.00	1.3398	3.798
12.17	.3444	.663
30.03	.3561	.164
37.40	.3594	2.951

Kohlrausch, 1879

%	$\varkappa$
18°	
5	648
10	1201
20	2075
30	2641
40	2502

Heydweiller, 1909

N	λ	N	λ
	18°		
0.1	94.02	2.0	66.5
0.2	90.3	4.0	52.9
0.5	82.6	5.5	44.5
1.0	76.0		

Okazaki, 1942

%	Verdet's constant (D)
25°	
0	0.013075
12.58	.01336
15.19	.01339
18.68	.01344
22.14	.01348
26.41	.01347
29.33	.01344

Jauch, 1921

N	U
18°	
0.5	0.9621
1	.9288
2	.8706
3	.8242
4	.7836

Lange and Eichler, 1927

mol% initial	mol% final	Q dil. by mole KF
23.4	21.9	+187.4
21.9	20.5	172.6
20.5	19.3	152.0
19.3	18.3	129.6
18.3	16.6	205.6
16.6	14.8	188.9
14.8	13.0	179.3
13.0	11.2	129.3
11.2	9.6	93.6
9.6	8.1	68.5
8.1	6.8	43.4
6.8	5.8	27.2
5.8	4.8	15.3
23.4	23.0	49.41
23.0	22.6	45.32
22.6	22.3	47.58
22.3	21.9	42.75
23.4	2.81	1618
23.1	2.67	1611
100	33.3	5912

Lange and Eichler, 1927

mol%	Q diss.* intermediate
0.179	+4106
.51	4064
.88	4054
1.0	4043
1.7	4021
2.4	3991
3.6	3970
4.7	3955

mol%	Q diss.* integral
0.36	+4106
0.72	4086
0.99	4075
1.4	4064
1.9	4056
2.9	4032
3.8	4017
5.0	4002

* Q diss., by mole KF

Kapustinskii and Ruzavin, 1955

%	$c.10^6$	%	$c.10^6$
	25°		
3.93	1446	17.5	1409
7.89	1434	24.8	1372
11.2	1425	25.0	1320

$c.10^6$= heat conductivity coefficient

Water (H_2O) + Potassium acid fluoride (KHF_2)

Jatlov and Polyakova, 1938

%	f.t.	%	f.t.
0	0	23.14	10
5	-2.9	28.15	20
10	-4.9	33.03	45
16.5	-7.6 E	44.08	60
19.70	0	53.28	80

Water (H_2O) + Potassium chloride (KCl)

Heterogeneous equilibria

a) Saturated vapour pressure

Wüllner, 1858

t	p 0%	p 9.1%	p 19.7%
23.1	21.01	20.31	19.62
27.1	26.66	25.76	24.77
29.1	29.95	28.91	28.01
31.5	34.36	33.17	31.78
37.3	48.25	46.46	44.58
39.4	53.14	51.16	49.07
41.5	59.48	57.20	54.82
44.0	67.79	64.91	62.04
47.3	80.33	76.82	73.40
49.1	87.93	84.46	80.99
51.7	100.00	95.54	90.89
55.1	118.04	113.32	108.70
56.5	126.24	121.05	115.56
58.5	138.76	133.22	127.08
62.5	166.98	159.67	152.55
64.5	182.12	174.92	167.32
65.6	201.07	193.07	184.67
68.9	222.29	213.11	204.42
70.5	238.24	228.38	218.52
72.5	260.60	249.05	237.90
75.4	293.30	281.08	269.05
78.7	336.33	322.63	308.35
81.5	376.85	361.78	346.50
83.4	406.57	389.53	372.67
86.0	450.34	431.79	412.53
88.7	500.00	479.50	457.97
92.6	579.67	556.21	531.48
94.9	631.44	604.89	577.47
100.3	768.20	737.42	706.30

Pauchon, 1880

t	p	t	p
19.60%		15.44%	
7.32	6.98	7.92	7.42
7.72	7.17	8.76	7.82
8.17	7.38	9.42	8.16
8.63	7.58	10.17	8.61
8.95	7.73	11.01	8.90
9.40	7.98	11.85	9.62
9.92	8.26	12.51	10.16
10.51	8.58	13.89	10.44
10.83	8.78	14.43	11.31
10.87	8.84	15.11	11.84
11.38	9.13	15.99	12.66
12.10	9.60	16.29	12.67
12.43	9.80	17.43	13.70
12.96	9.91	17.98	14.19
13.31	10.20	18.71	14.89
13.72	10.67	19.35	15.42
14.23	10.98	20.01	16.09
14.88	11.40	20.73	16.87
15.51	11.91	21.16	17.28
16.00	12.39	21.87	18.09
17.21	13.22	22.31	18.51
17.85	13.75	22.97	19.32

18.35	14.18	23.41	19.80
18.70	14.55	24.13	20.67
19.10	14.86	24.83	21.59
19.81	15.45	25.32	22.22
20.45	16.09	25.99	23.07
21.17	16.74	26.72	24.12
22.92	18.66	27.39	25.09
23.71	19.59	28.02	26.01
24.31	20.37	28.85	27.32
24.99	21.46	29.45	28.27
25.41	21.48	29.65	28.61
26.19	22.72	30.37	29.74
27.01	23.92	31.02	30.84
27.31	24.40	31.83	31.26
27.92	25.24	32.51	33.63
28.43	26.49	33.28	35.12
28.85	26.64	33.92	36.46
29.30	27.35	34.41	37.46
30.03	28.53		
30.76	29.74		
31.85	31.60		
32.54	32.87		
33.09	33.91		
34.11	35.88		
34.93	37.60		
35.00	37.77		

10.56%		5.89%	
6.00	6.73	7.11	7.37
6.38	6.89	7.42	7.55
6.68	7.03	7.83	7.75
6.89	7.12	8.18	7.90
7.27	7.30	9.10	8.40
7.82	7.58	9.92	8.92
8.21	7.82	10.71	9.41
8.59	8.00	10.99	9.60
8.98	8.23	11.68	10.04
9.27	8.36	12.81	10.73
9.65	8.59	13.30	11.09
10.33	9.02	14.57	12.06
10.41	9.01	15.66	12.90
11.17	9.51	16.29	13.48
11.62	9.98	17.41	14.48
12.16	10.12	18.59	15.49
12.60	10.46	20.33	17.24
13.14	11.19	21.30	18.36
13.83	11.36	22.43	19.71
14.21	11.54	23.69	21.26
14.72	11.96	24.85	22.77
15.20	12.32	25.83	24.13
15.73	12.75	26.62	25.30
16.26	13.21	27.15	26.10
16.94	13.83	28.68	28.54
17.63	14.35	29.13	29.30
18.16	14.89	29.57	30.03
19.92	16.64	29.99	30.79
19.96	16.65	30.33	31.51
20.59	17.26	30.69	32.06
21.35	18.10	31.01	32.64
22.49	19.44	31.41	33.40
23.10	20.10	31.84	34.23
23.56	20.69	32.23	34.99
23.71	20.89	33.58	37.76
24.47	21.90	34.86	40.60
25.38	23.11	35.21	41.34
25.52	23.24		
27.48	26.05		
28.88	28.23		
29.71	29.75		
30.80	31.70		
31.76	33.48		
32.28	34.51		
34.42	38.86		

Tammann, 1885

t	p			
	0%	12.14%	21.10%	21.47%
45.56	74.0	69.7	65.4	65.3
66.21	198.0	185.7	175.0	174.0
71.51	249.5	234.6	220.0	218.6
81.55	378.3	354.9	332.8	331.5
87.81	483.8	454.8	425.3	423.8
91.35	553.5	519.3	487.5	485.2
98.05	708.7	664.3	624.9	622.2
99.43	758.0	712.7	668.2	665.0

Tammann, 1885

%	p	%	p
100°			
2.41	751.9	21.53	667.9
4.62	744.8	24.06	652.0
10.81	720.3	27.19	631.4
16.18	696.8	33.85	589.3

Emden, 1887

t	p	t	p	t	p
9.13%		16.7%		23.08%	
21.68	18.7	14.18	11.15	15.50	11.65
26.82	25.6	25.59	22.15	19.30	14.75
28.68	28.3	29.69	28.4	24.95	20.5
34.77	40.0	33.69	35.8	30.36	28.15
40.20	53.3	40.46	51.3	34.58	35.75
45.39	70.05	44.26	62.5	39.88	47.45
49.02	83.7	49.67	83.1	45.57	63.8
55.62	115.35	52.98	96.9	51.17	84.4
59.78	140.2	60.93	140.7	55.26	103.0
64.84	176.2	64.10	163.9	60.00	128.9
68.42	208.0	68.92	201.05	65.74	167.3
74.68	271.8	77.18	287.5	70.30	204.2
80.00	338.1	80.05	323.5	76.68	268.15
84.30	410.5	85.35	399.3	80.68	315.0
90.02	502.3	90.09	478.7	85.08	376.1
94.88	603.3	94.74	570.1	90.94	471.5
				95.00	548.1

Dieterici, 1891

%	p
0°	
0	4.620
6.93	.472
12.97	.326
18.27	.190
22.18	.083

Van't Hoff and Meyerhoffer, 1901

8.1 mol% (25°) p= 19.2 mm

Speranski, 1910

t	p	t	p
sat.sol.			
23.2	18.58	38.36	41.79
26.3	22.09	41.87	50.10
29.34	25.88	45.11	58.65
32.21	30.28	47.89	67.36
35.7	36.50	51.1	78.6

Krauskopf, 1910

%	p
40°	
7.98	52.95
11.74	51.79
16.38	50.21
18.33	49.55
22.20	48.10

Perman and Price, 1912

t	c	p	t	c	p
70°			90°		
69.95	4.89	231.3	90.02	4.18	514.5
70.00	7.33	229.6	90.00	7.23	509.7
69.99	9.77	225.1	89.98	11.58	498.28
70.01	14.66	219.0	89.99	13.19	494.3
70.00	19.52	211.7	90.03	31.76	441.88
70.02	24.55	206.1	89.96	32.39	438.76
70.00	29.53	202.1			

Edgar and Swan, 1922

t	p	t	p
sat.sol.			
10	13.97	25	20.20
20	14.97	26	21.33
21	15.98	27	22.64
22	17.01	28	23.99
23	18.05	29	25.37
24	19.10	30	26.88

Pohle, 1927

t	p	t	p
sat.sol.			
30	28	70	187
35	36	75	272
40	46	80	283
45	58.5	85	342
50	78	90	406
55	98	95	490
60	123	100	584
65	152	105	688

Harrison and Perman, 1927

%	p	%	p
40°		50°	
0	55.2	0	92.4
3.16	55.0	4.23	90.1
3.90	54.9	8.81	89.1
5.82	54.8	10.91	87.6
8.68	53.6	14.27	86.3
9.77	53.1	21.46	81.7
10.43	52.9	26.32	77.9
13.03	51.7	30.00	74.7
17.84	50.3	70°	
18.80	50.1		
22.13	48.2	0	233.8
23.41	48.4	4.14	230.2
27.30	47.0	5.85	229.5
29.44	46.1	7.16	228.6
60°		8.74	226.4
		10.18	224.4
0	149.3	12.35	221.2
3.80	147.1	13.66	219.3
5.15	146.4	17.11	214.4
8.08	144.3	21.92	205.6
10.91	142.1	27.00	195.0
13.27	139.8	28.38	190.8
15.92	137.0	32.60	184.6
21.61	131.4	80°	
24.64	128.4		
28.00	123.7	0	355.4
31.30	120.3	9.45	341.6
		11.47	338.1
		15.13	330.5
		18.59	322.3
		22.73	311.3
		27.24	297.2
		30.01	289.1
		34.00	277.4

Leopold and Johnston, 1927

t	m	p	
		0%	sol.
21.42	4.664	19.14	16.23
25.62	.849	24.65	20.72
27.56	.926	27.65	23.27
33.64	5.178	39.11	32.46
39.31	.402	53.32	43.70
44.03	.576	68.37	55.62

Pearce and Snow, 1927 and Pearce and Nelson, 1932

m	p	m	p
25°			
0	23.752	2.00	22.243
0.20	23.597	2.50	21.877
.40	23.448	3.00	21.497
.60	23.296	3.50	21.089
.80	23.155	4.00	20.696
1.00	23.017	4.81	20.021
1.50	22.620		

Adams and Merz, 1929

t	p	t	p
sat.sol.			
10	8.07	25	19.89
15	11.05	30	26.75
20	15.04	40	44.99
		50	73.97

Ebert, 1930

%	p 0°	5.2°	20°	30°
0	4.58	6.64	17.54	31.83
5	.50	.55	17.15	31.10
10	.35	.35	16.65	30.10
15	.20	.10	16.10	29.30
20	.05	5.90	15.60	28.40
25.5	-	-	14.95	-
(satd.)				

Lannung, 1934

m	p	m	p
18°			
0.3380	15.31	1.761	14.65
.4054	15.27	2.778	14.12
.5290	15.22	3.035	13.98
.8740	15.05	4.340	13.30
1.0030	15.00	satd.	13.20
1.5280	14.74		

Shibata and Niwa, 1935

t	p	t	p
sat.sol.			
0	4.067	5.9	6.095
1.1	4.387	6.9	6.520
2.0	4.666	7.92	6.999
3.0	5.001	8.9	7.472
4.0	5.355	9.8	7.930
4.93	5.714		
4.0250 N			
5.9	6.104	7.9	7.025
6.9	6.539	9.8	7.995
3.9078 N			
5.0	5.764	7.9	7.046
6.0	6.184	8.9	7.547
6.9	6.563		
3.7033 N			
1	4.332	7	6.615
4	5.394	10	8.106
3.3325 N			
1	4.414	7	6.722
4	5.479	10	8.106
1.8859 N			
1	4.643	7	7.077
4	5.766	10	8.675
1.5790 N			
1	4.690	7	7.148
4	5.824	10	8.763
0.9250 N			
1	4.784	7	7.294
4	5.943	10	8.944
0.4077 N			
1	4.866	7	7.419
4	6.044	10	9.096
0.3335 N			
1	4.878	7	7.438
4	6.059	10	9.120
0.2849 N			
1	4.882	7	7.445
4	6.064	10	9.129
0.2786 N			
1	4.886	7	7.450
4	6.068	10	9.136

t = T - 273.100

Keevil, 1942

mol%	t	p	mol%	t	p
16.0	190.0	8.58	32.3	428.6	150.7
16.8	213.0	14.50	33.5	439.5	160.4
18.0	237.2	19.45	35.4	457.2	187.0
18.0	237.6	19.81	37.0	472.0	185.0
19.8	269.2	32.33	37.7	480.4	190.7
21.6	298.0	47.80	38.1	485.8	194.0
21.6	298.1	48.26	39.3	493.8	200.1
23.8	330.0	68.40	39.5	497.2	202.0
25.2	349.4	82.80	41.6	516.1	214.0
27.0	371.0	100.50	42.7	526.4	218.5
27.2	372.5	101.80	43.2	529.7	217.0
27.4	375.0	102.20	44.4	538.9	218.0
29.3	397.3	121.80	50.9	584.5	220.0
32.2	427.2	151.00	62.4	645.0	197.0

Morey and Chen, 1956

t	P	t	P
sat.sol.			
374	110.5	600	215.5
400	127.5	700	195.5
500	206.5		

Foote, Saxton and Dixon, 1932

$$\lg P = \frac{2995.5}{T} - 6.680 \lg T + 0.001024\, T + 27.569$$

Smith, Combs and Googin, 1954

m	Dp/p_o
30.01°	
0.754	-0.0225
1.081	.0341
1.665	.0559
1.869	.0593
2.629	.0845
2.824	.0925
3.957	.130

Boswell and Cantelo, 1922

N	Dp/p_o
23°	
4.600	-0.147
3.000	.090
2.000	.064
1.000	.030

Hartung, 1920

%	p_1-p sat.sol	%	p_1-p sat.sol.
30.01°			
0	5.25	9.63	3.84
0.70	5.19	10.48	3.66
0.91	5.14	14.47	2.98
1.00	5.09	14.83	2.91
2.00	4.98	18.67	1.13
2.34	4.95	20.63	1.61
2.66	4.88	22.48	1.26
3.15	4.79	25.37	0.47
5.52	4.44	27.22	sat.sol.
8.26	4.06		

Nicol, 1886

mol%	Dp 70°	75°	80°	85°	90°	95°
2	-7.6	-9.6	-12.2	-14.7	-17.4	-22.3
3.9	16.6	20.5	25.1	31.0	37.4	46.0
5.7	25.3	31.2	38.9	47.6	57.8	70.6
7.4	34.5	42.7	53.2	64.9	74.7	95.1
9.1	44.3	54.7	67.5	82.2	99.7	120.7

Witt, 1900

mol %	Dp	mol %	Dp
2.41	-0.60	4.70	-1.25
3.22	-0.99	6.87	-1.91

Lovelace, Frazer and Miller, 1916

m	Dp	m	Dp
20°			
0.2	-0.110	1.0	-0.547
.4	.217	.2	.663
.6	.329	.5	.826
.8	.438	2.0	1.102
.9	.493		

Lovelace, Frazer and Seave, 1921

m	Dp	m	Dp
20°			
0.0498	-0.0293	0.9968	-0.5525
.0704	.0408	1.2461	.6904
.0997	.0573	.5094	.8383
.2013	.1136	.7072	.9482
.2985	.1672	.9938	1.1077
.3855	.2144	2.2450	.2492
.4973	.2770	.4916	.3893
.6000	.3347	3.0017	2.0203
.6995	.3899	4.0070	.2741
.7958	.4412		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
25°			
0.1	0.927	1.2	0.899
.2	.913	.4	.901
.3	.906	.6	.904
.4	.902	.8	.908
.5	.899	2.0	.912
.6	.898	2.5	.924
.7	.897	3.0	.937
.9	.897	3.5	.950
1.0	.897	4.0	.965
		4.5	.980

Brown and Delaney, 1954

m	osmotic coefficient	m	osmotic coefficient
25°			
0.02515	0.9622	0.3180	0.9009
0.04445	0.9515	0.4254	0.8952
0.05621	0.9451	0.6025	0.8946
0.08020	0.9320	0.7741	0.8920
0.1210	0.9193	1.8631	0.9096
0.1901	0.9053	2.3783	0.9205

b) Boiling curve

Kremers, 1856

Sat.sol.: b.t.=110°

De Heen, 1881

%	b.t.
23.58	104.8
15.70	103.0
10.49	101.5

Gerlach, 1886

%	b.t.	%	b.t.
0	100	24.87	104.5
4.67	100.5	26.58	105
8.43	101	28.21	105.5
11.58	101.5	29.78	106
14.31	102	31.27	106.5
16.74	102.5	32.60	107
18.95	103	33.99	107.5
21.05	103.5	35.27	108
23.02	104	36.47	108.5

Roloff, 1893

%	b.t.	%	b.t.
0	99.756	7.41	100.720
0.5	99.812	8.26	100.848
0.99	99.860	9.09	100.979
1.47	99.921	9.91	101.112
1.96	99.970	10.71	101.245
2.91	100.099	11.50	101.279
3.85	100.223	12.28	101.513
4.76	100.345	13.04	101.648
5.66	100.468	14.27	101.877
6.54	100.594	15.57	102.132

Baroni, 1893

%	b.t.	%	b.t.
1.98	100.261	10.74	101.523
3.96	100.515	13.33	101.981
6.16	100.813	16.72	102.634
8.15	101.114	19.75	103.278

Berkeley and Applebey, 1911

Sat.sol.: b.t.= 108.599

Buchanan, 1899

%	b.t.	%	b.t.	%	b.t.
		772.0 mm			
36.20	108.83	28.57	105.98	21.37	104.00
36.55	108.72	26.85	105.49	19.20	103.50
34.21	108.02	25.19	105.00	16.91	103.02
31.40	106.97	23.38	104.50	0.00	100.44
29.95	106.46				
		763.3-767.7 mm			
56.13	108.43	49.11	106.95	43.80	105.47
56.32	108.32	47.45	106.44	34.76	104.97
50.64	107.47	45.79	105.96	0.00	100.10
		738.8 mm			
35.97	107.51	26.99	104.26	18.99	102.23
35.80	107.31	25.21	103.75	16.46	101.72
33.23	106.79	23.29	103.25	0.00	99.21
30.22	105.27	21.23	102.75		
		619.2 mm			
35.95	102.24	26.68	99.15	17.90	97.12
33.19	101.19	20.41	97.63	0.00	94.37
30.11	100.16				
		613.7-614.0 mm			
35.42	102.00	25.21	98.55	18.57	97.03
35.38	101.80	23.26	98.04	16.26	96.53
35.30	100.06	21.14	97.54	0.00	94.12
35.26	99.05				
		550.4 mm			
35.49	98.89	29.92	96.87	17.68	93.84
35.53	98.70	26.42	95.80	0.00	91.20
32.92	97.89	22.37	94.85		
		549.8 mm			
35.85	98.87	30.05	96.89	26.59	95.87
35.25	98.61	28.60	96.38	0.00	91.18
33.14	97.90				
		non sat. sol.			
20.80	103.34	9.89	101.12	5.08	100.41
16.41	102.33	8.55	100.91	4.30	100.31
13.92	101.82	7.23	100.71	3.38	100.21
11.09	101.32	5.83	100.51	0.00	99.75

Kahlenberg, 1901

%	b.t.	%	b.t.	%	b.t.
		737.6 mm			
2.07	100.293	12.27	102.004	21.72	104.266
4.87	100.700	16.79	102.975	24.15	105.009
8.27	101.247	19.45	103.653		
		744 mm			
1.88	100.255	10.41	101.625	20.71	103.911
3.90	100.554	14.17	102.406	23.30	104.756
6.74	100.986	17.59	103.206		
		736.1 mm			
5.36	100.65	16.44	102.60	27.80	105.60
6.91	100.91	19.30	103.18	29.84	106.28
9.09	101.25	21.37	103.75	31.42	106.88
11.07	101.59	23.01	104.30	32.86	107.60
13.75	102.05	25.88	105.00		

c) Freezing curve .

Rüdorff, 1861

%	f.t.	%	f.t.
0.99	-0.45	7.41	-3.55
1.96	-0.90	9.09	-4.40
3.85	-1.80	10.71	-5.35
5.66	-2.65		

de Coppet, 1872

%	f.t.	%	f.t.
5.98	-2.87	14.49	-7.75
9.58	-4.80	17.00	-9.70
11.30	-5.85	20.31	-11.50

Roloff, 1895

%	f.t.	%	f.t.
0	0	12.68	-6.37
0.84	-0.384	14.40	-7.34
1.78	-0.827	16.62	-8.66
2.99	-1.389	18.86	-10.04
5.76	-2.721	19.45	-10.46
7.25	-3.417	19.75	-10.75
9.70	-4.696		

Van't Hoff and Meyerhoffer, 1901

5.6 mol% E: -11.1°

Jones, 1904

N	f.t.	N	f.t.
0.20	-0.685	2.00	-6.944
0.50	-1.692	3.00	-1.062
1.00	-3.400		

Meusser, 1905

%	f.t.	%	f.t.
12.5	-5.5	20.6	-4.5
13.6	-6	21.4	-1
14.9	-7	22.2	+2.5
16.7	-8	23.0	+7.5
17.7	-8.5	23.8	+11.5
19.3	-9	25.0	+18.5

Jones and Stine, 1908

N	f.t.	N	f.t.
0.45	-1.52	1.2972	-4.48
0.65	-2.24	1.3	-4.49
0.85	-2.91	1.7	-5.93
0.90	-3.08	2.1	-7.43
1.05	-3.58	2.5945	-9.29

Rodebush, 1918

%	f.t.
6.62	-3.07
9.72	-4.66
14.80	-7.51
18.49	-9.84
19.22	-10.34
19.74	-10.66 E

Jones and Bury, 1927

m	f.t.	m	f.t.
0.0988	-0.340	0.8303	-2.719
.1205	.412	.9182	3.001
.1246	.425	1.020	.328
.1816	.616	.112	.618
.2381	.803	.220	.960
.2710	.913	.304	4.230
.2998	1.007	.397	.528
.4042	.347	.484	.804
.4549	.511	.608	5.208
.5204	.723	.615	.223
.6047	.992	.704	.519
.7013	2.303	.778	.757
.7856	.577	.818	.882

Barnes and Maass, 1930

%	f.t.	%	f.t.
4.95	-2.24	18.94	-9.78
9.48	-4.60	19.02	-10.31
13.70	-6.88	19.49	-10.36
17.85	-9.48		

Rolla and Mariani, 1948

N	f.t.	N	f.t.
0.00645	-0.0234	0.50444	-1.6927
.01621	-0.0580	.61645	-2.0667
.03814	-0.1342	.67218	-2.2515
.07660	-0.2659	.76296	-2.5545
.13511	-0.4638	.85690	-2.8692
.20174	-0.6865	.95716	-3.1949
.28415	-0.9603	1.12759	-3.7817
.40133	-1.3497	.20551	-4.0479

Fialkov and Chernogorenko, 1955

%	f.t.	E	%	f.t.	E
0.47	-0.21	-10.76	13.3	-6.68	-10.76
0.85	-0.39	"	15.3	-7.85	"
1.76	-0.81	"	17.0	-8.88	"
3.51	-1.60	"	18.4	-9.77	"
6.90	-3.30	"	18.9	-10.17	"
7.12	-3.40	"	18.9	-9.80 II	-
11.1	-5.50	"	20.0	-9.80	-10.76

Mun and Darer, 1956

m	f.t.
0.64	-2.10
1.11	-3.50
1.80	-5.80
2.46	-7.90

d) Solubility curve

Kopp, 1840

%	f.t.
25.17	11.8
25.87	13.8
25.93	15.6

Mulder, 1864

%	f.t.	%	f.t.
22.18	0	27.28	29.75
22.78	4.25	29.48	45
25.20	16.5	36.91	b.t.107.65
26.42	25		

Chernogorenko, 1956

E : -10.8

Gérardin, 1865

%	f.t.
25.15	18
27.43	30
28.62	40
31.03	57

de Coppet, 1883

%	f.t.	%	f.t.
19.65	-11.0	28.92	41.45
20.48	-6.4	29.75	46.15
21.82	0.0	29.99	48.8
22.70	3.9	30.79	55.1
23.58	9.4	31.46	60.55
24.35	11.4	32.05	64.95
24.62	14.95	32.78	71.65
25.55	19.0	33.00	74.25
26.52	25.7	33.88	80.75
27.18	29.25	34.43	86.6
28.42	38.0	34.90	91.4

Andreae, 1884

%	f.t.	%	f.t.
21.87	0	27.16	30
22.68	4	28.63	40
23.81	10	30.00	50
25.54	20	31.26	60

Berkeley, 1904

%	f.t.	%	f.t.
	sat.sol.		
22.05	0.70	33.15	74.80
25.58	19.55	34.76	89.45
27.71	32.80	36.75	108.00
31.43	59.85		

Chernai, 1912

%	f.t.	%	f.t.
	sat.sol.		
32.57	70	25.43	20
31.26	60	23.80	10
30.00	50	22.18	0
28.63	40	19.99	-10
27.16	30		

Sudhaus, 1914

mol %	f.t.	mol %	f.t.
7.58	19.3	8.88	40.1
8.25	29.7	9.41	54.5

Ceitlin, 1925

N	f.t.
3.71	10.2
3.90	20.0

Cornec and Hering, 1926

%	f.t.
21.83	0
26.43	25
30.07	50
33.16	75
35.70	100

Scott and Frazier, 1927

26.38% f.t.= 25°

Buchanan , 1912 - 13

m	t
4.7619	23.4

Tilden and Shenstone, 1885

%	f.t.
37.36	125
40.94	133
41.45	144
42.92	175
43.66	180

Etard, 1894

%	f.t.	%	f.t.
38.6	142	43.2	190
38.8	150	42.9	200
41.2	175	47.6	242
41.8	180	100.0	732

Froehlich, 1929

%	f.t.	%	f.t.
38.30	130	41.40	150
38.40	130	43.20	178
40.20	131	43.30	180
40.80	133		

Cornec and Krombach, 1932

%	f.t.	%	f.t.
21.92	0	36.50	108.5
25.57	20	37.65	120
28.65	40	39.60	140
31.29	60	42.42	169.5
33.59	80	44.34	189.6
35.69	100		

Akhumov and Vasiliev, 1932

%	f.t.	%	f.t.
0	0	43.82	190
19.35	-10.66 E	44.50	200
21.88	0	45.41	210
28.43	33.3	45.65	220
31.97	66.6	45.89	230
35.9	100	48.00	240
36.71	110	48.82	250
37.53	120	49.49	260
38.60	130	50.12	270
39.29	140	50.36	280
40.48	150	51.73	290
41.12	160	52.13	300
42.20	170		

Denecke, 1919

P Kg	f.t.	
	19.8%	
1	-11.1	KCl + ice I + sol.
1326	23.7	KCl + ice I + sol.
1948	31.6	KCl + ice I + sol.
2815	31.7	KCl + ice III + sol.
3007	30.7	KCl + ice III + sol.
3015	30.6	KCl + ice III + sol.
2210	35.1	KCl + ice I + ice III
2465	40.0	tr.t. KCl + ice III + ice I
2764	38.0	" " " " "
2910	37.5	" " " " "

Jones (B.M.) and Shah, 1913

%	t. spontan. cryst.	%	t. spontan. cryst.
0	-0.6	21.08	-13.5
1.96	-2.4	21.84	-10.2
3.84	-3.6	22.43	-7.4
3.96	-3.8	22.53	-6.5
5.69	-5.1	24.15	+1.2
7.40	-6.0	24.55	4.3
10.49	-8.1	24.66	4.7
13.40	-10.5	26.23	13.3
15.80	-13.0	27.88	23.6
16.94	-13.5	28.73	29.9
18.99	-14.4	30.04	38.2
19.41	-15.1	30.55	43.2
19.55	-15.3	30.56	43.2
19.97	-15.5	31.80	52.2
20.12	-15.7	32.58	60.0
20.32	-15.7	33.48	65.0
20.34	-15.5	33.78	67.9
20.55	-15.2	35.10	81.9
20.61	-14.9		

Akhumov and Pylkova, 1956

%	f.t.	t spontan.cryst.
33.70	79	56
35.96	101	77
37.79	119	94
38.30	125	99
39.95	142	112
43.12	177	148
44.72	195	168
46.79	219	193
49.55	250	216

Properties of phases .

a) Density .

Stocker, 1920

%	t	d
0	17.0	0.9988
6.73	18.4	1.0423
12.65	19.1	.0819
21.06	18.3	.1416

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0111	16	1.1046
4	.0240	18	.1188
6	.0370	20	.1331
8	.0501	22	.1478
10	.0635	24	.1646
12	.0771	25	.1702
14	.0906	25.44	.1736

Schiff, 1858, 1859 and 1860

%	d	%	d
	17.5°		
0	0.9987	13	1.0852
1	1.0049	14	.0923
2	.0112	15	.0994
3	.0176	16	.1066
4	.0241	17	.1138
5	.0306	18	.1210
6	.0372	19	.1283
7	.0438	20	.1357
8	.0504	21	.1431
9	.0572	22	.1506
10	.0641	23	.1582
11	.0711	24	.1658
12	.0781		

Gerlach, 1859

%	d	%	d
	15°		
0	0.99913	13	1.08559
1	1.00563	14	.09250
2	.01212	15	.09940
3	.01861	16	.10654
4	.02511	17	.11368
5	.03160	18	.12081
6	.03825	19	.12796
7	.04491	20	.13509
8	.05156	21	.14248
9	.05822	22	.14987
10	.06487	23	.15727
11	.07177	24	.16466
12	.07868		

Fouqué, 1867

%	d	
	0°	21°
0.75	1.0050	1.0039
3.62	1.0227	1.0207
7.15	1.0476	1.0442
19.63	1.1382	1.1310

Kohlrausch and Grotrian, 1875

%	d
	18°
5	1.0309
10	.0639
15	.0978
21	.1410

Kohlrausch, 1879

%	d
	18°
5	1.0308
10	.0638
15	.0978
20	.1335
25	.1408

Rönnberg, 1980

t	d				
	5%	10%	15%	20%	25%
4	1.03104	1.06017	1.08745	1.11328	1.13773
5	.03090	.05998	.08722	.11300	.13742
10	.03022	.05902	.08606	.11160	.13587
15	.02928	.05781	.08463	.10998	.13409
20	.02812	.05640	.08301	.10819	.13216
25	.02672	.05477	.08119	.10622	.13006

De Heen, 1881

t	d	t	d	t	d
		23.58 %			
10.00	1.000000	21.19	0.996255	39.86	0.989099
15.23	0.998317	28.45	0.993592	48.80	0.985267
		15.70 %			
10.00	1.000000	31.24	0.993019	57.53	0.981695
16.97	0.997964	39.53	0.989722	65.05	0.977921
24.42	0.995469	49.62	0.985847	70.60	0.974978
		10.49 %			
10.00	1.000000	30.24	0.993746	54.03	0.983799
15.20	0.998556	39.42	0.990214	60.70	0.980469
22.83	0.996232	48.07	0.986478	70.47	0.975605

Volkmann, 1882

%	d
15-16°	
24.66	1.1696
15.21	1.1011
7.19	1.0463

Nicol, 1883

mol%	d 20°	d 40°	mol%	d 20°	d 40°
4.8	1.11248	1.10414	1.0	1.02386	-
3.9	.09221	-	0.5	.01131	1.01712
2.0	.04773	1.04050	0	0.99823	0.99224

Bender, 1883

N	d
15°	
1	1.0444
2	1.0887
3	1.1317

Rother, 1884

%	d	%	d
	15°		
2.50	1.01548	14.84	1.09823
4.95	1.03133	19.62	1.13496
7.45	1.04788	23.04	1.15739
10.01	1.06492		

Nicol, 1884

t	d
8.5 N (20°)	
20	1.1806
40	.16460
98	.09287

Thomsen, 1886

mol%	d	mol%	d
	18°		
6.3	1.1468	1.0	1.0258
3.2	.0800	0.5	.0136
1.9	.0496	0	0.9986

Röntgen and Schneider, 1888

%	d	%	d
	18°		
0.00	0.9987	9.82	0.0631
0.871	1.0043	15.03	0.0983
2.48	1.0146	19.94	0.1334
4.88	1.0300	25.52	0.1747
7.51	1.0481		

Schumann, 1887

%	d	%	d
	15°		
0	0.9991	10.68	1.0696
2.52	1.0155	16.81	.1123
5.35	.0337	22.83	.1559

Drecker, 1888

%	d 0°	t	d	$\tau.10^6$ 20°
2.49	1.0172	16.90	1.01576	228
4.40	.0338	13.84	.0320	246
8.28	.0573	15.00	.0543	278
13.02	.0878	"	.0867	308
16.75	.1569	"	.1129	327
24.31	.1733	"	.1680	359

Le Blanc, 1889

%	d
20°	
0	0.99823
5.05	1.03089
20.55	1.13731

Bender, 1884

M	d	$\tau.10^6$	
	15°	15-20°	20-25°
4.23	-	-	1718
4.11	1.1762 a	1625	1712
3.92	.1696	1606	-
3.5	.1523	1577	1672
3	.1317	1521	1620
2.5	.1086	1440	1566
2	.0887	1360	1508
1.5	.0672	1255	1427
1	.0444	1109	1312
0.75	.0339	1043	1248
0.375	.0163	897	1135
0.185	.0081	823	1074
0	0.9991	738	1011

a= sat.sol. at 14.9°

Traube, 1885

c	%	d
	15°	
10	9.41	1.0623
20	17.85	1.1205

Brückner, 1891

mol%	d	mol%	d
	15°		
0	0.9992	2	1.0899
0.5	1.0225	2.5	.1116
1	.0457	3	.1332
1.5	.0680	3.5	.1540

Bodländer, 1891

%	d	%	d
14.5°		17.5°	
0	0.9992	0	0.9987
24.83	1.1710	25.06	1.1720

Charpy, 1893

%	d	%	d
	0°		
4.4968	1.0307	14.4707	1.1017
7.5440	.0520	17.7214	.1257
11.0757	.0768	20.7840	.1488

Sentis, 1897

mol %	t	d	mol %	t	d
1	19.4	1.0236	3	16.3	1.0743
1	16.45	1.0249	5	16.4	1.1204
1	12.2	1.0258	5	21.8	1.1162
2	12.3	1.0512	6	21.2	1.1378
2	16.55	1.0502			

Gilbault, 1897

%	d	%	d
	20°		
0	0.9982	15	1.0984
5	1.0307	20	.1321
10	1.0639	23.97	.1633

Barnes and Scott, 1898

%	d	%	d
	20.1°		
0	0.9982	13.17	1.0617
1.197	1.0040	18.06	.0866
3.456	.0148	24.79	.1215
4.419	.0193	31.12	.1554
6.610	.0297	36.43	.1853
8.412	.0386		

Oppenheimer, 1898

%	d
20°	
6.64	1.0424
12.82	.0829
17.66	.1166
23.93	.1617

Linebarger, 1899

%	d
25°	
9.09	1.0555
10.71	.0665
16.67	.1069
23.08	.1526
26.44	.17885

Barnes and Johnson, 1902

%	d	%	d
	25.5°		
0	0.9969	11.87	1.0520
0.97	.9996	14.20	.0646
1.61	1.0027	18.81	.0862
3.10	.0098	21.92	.1022
5.50	.0215	23.91	.1130
7.72	.0316	29.64	.1431
9.50	.0403		

Taylor and Rankin, 1903-04

N	d	N	d	N	d
0°		15°		25°	
1	1.0480	1	1.0455	1	1.0433
2	.0935	2	.0901	2	.0877
3	.1371	3	.1333	3	.1295

Grabowsky, 1904

%	d 10°	d 18°	d 30°
0	0.999731	0.99863	0.99567
9.47	1.0627	1.0603	1.0566
18.27	.1248	.1212	.1159
23.32	.1618	.1585	.1535

Berkeley, 1904

t	d	t	d
	sat.sol.		
0.70	1.1540	74.80	1.2032
19.55	.1738	89.45	.2069
32.80	.1839	108.00	.2118
59.85	.1980		

Agerer, 1905

%	d
18°	
7.505	1.0482
11.958	.0781
17.098	.1138

Bousfield, 1905

%	d 15°	d 18°	d 21°
21.000	1.14214	1.14094	1.13970
20.000	.13486	.13375	.13255
19.000	.12759	.12651	.12534
18.000	.12050	.11935	.11817
17.000	.11334	.11228	.11115
15.0095	.09924	.09820	.09712
12.9993	.08530	.08434	.08334
12.001	.07847	.07753	.07654
11.4589	.07478	.07387	.07290
10.0001	.06494	.06407	.063095
9.0026	.05815	.05732	.05639
7.1383	.04567	.04492	.04404
5.9998	.03823	.03745	.03663
4.00016	.02507	.02437	.02359
3.0003	.01857	.01794	.01720
2.000	.01207	.011475	.01077
1.00005	.00562	.00507	.00439
0.000	0.99915	0.99866	0.99807

Zecchini, 1905

%	t	d	%	t	d
19.769	8.6	1.13469	12.422	7.7	1.08286
17.628	7.4	.11945	5.805	8.8	1.03786
15.359	6.9	.10356	0	8.0	0.99988

Dinkhauser, 1905

%	d
18°	
2.290	1.0135
5.843	.0370
8.717	.0560
12.481	.0813
18.521	.1238

Getmann, 1907

%	d
18°	
4.96	1.10199
9.44	.0584
13.68	.0970
17.61	.1355
21.77	.1484
25.79	.1614

Chéneveau, 1907

%	d	%	d
	15°		
24.26	1.1676	13.06	1.0867
22.14	.1515	10.61	.0699
19.97	.1361	8.10	.0530
17.74	.1194	5.49	.0357
15.42	.1031	2.77	.0180
		0	0.9991

%	t	d
23.57	18.7	1.1609
13.99	19.1	.0907
4.97	19.1	.0303

Getman and Wilson, 1908

%	d	%	d
	20°		
0	0.99823	15	1.10036
5	1.03250	20	1.13608
10	1.06580	24	1.16568

Jones and Stine, 1908

N	d	N	d
	0°		
0.45	1.000525	1.3	1.060508
0.65	.009167	1.7	.07712
0.85	.018038	2.1	.097828
0.90	.039744	2.5945	.121016

Jaquerod, 1909

t	d	t	d
	70.4g/100cc		
0	1.0457	54.4	1.0280
14.9	1.0429	78.6	1.0192
25.5	1.0399		
	140.8g/100 cc		
0	1.0887	54.5	1.0684
15.2	1.0851	78.7	1.0550
25.0	1.0815		
	211.2g/100 cc		
0	1.1308	54.3	1.1084
14.8	1.1256	78.6	1.0952
25.4	1.1216		
	281.6g/100 cc		
0	1.1703	55.2	1.1462
17.5	1.1644	78.2	1.1335
25.0	1.1611		

Guerdjikowa, 1910

%	d
25°	
10.345	1.0724
19.143	.1240
25.830	.1833

Baxter, Boylston and al., 1911

%	d	%	d
	25°		
0	0.99707	7.1573	1.0428
0.9466	1.00303	7.1699	.0430
.9716	.0031	9.0221	.0549
.9769	.0032	9.4153	.0577
.9983	.00337	9.5004	.05814
1.0013	.0036	13.758	.0865
2.4297	.0122	13.755	.0872
2.4623	.0124	17.852	.1155
4.6474	.02654	17.898	.1160
4.8150	.0285	21.522	.1426
4.8318	.0281	21.770	.1441
4.8340	.0282		

Rubien, 1911

N	d	N	d
	18°		
0	0.9986	1	1.0448
0.1	1.00340	2	.0887
0.2	.00813	3	.1316
0.5	.02217		

Chernai, 1912

%	t	d	%	t	d
		sat.sol.			
19.99	-10	1.139	28.63	+40	1.190
22.18	0	.156	30.00	50	.195
23.80	+10	.168	31.26	60	.199
25.43	20	.177	32.57	70	.203
27.16	30	.183			

Clausen, 1912

N	t	d	N	t	d
0.498	18.0	1.02201	2.013	18.1	1.08930
	29.9	.01870		29.9	.08491
1.005	18.0	.04511	3.004	18.1	.13198
	29.9	.04133		29.9	.12710

Buchanan, 1912-13

N	d 19.5°	d 15°
0.5	1.021312	1.022321
.25	.010023	.011063
.12	.004251	.005080
.06	.001340	.002123
.03	0.999859	.000659
.015	.999112	0.999888
.008	.998736	.999538
.004	.998565	.999126
.002	.998454	-

m	d 19.5°	m	d 23.0°
16.94	1.0432	0.06	1.000531
212.98	.0835		
318.29	.1204		
422.98	.1543		

m	t	d
	sat.sol.	
4.7619	23.4	1.1768

Baxter and Wallace, 1916

%	d	%	d
	100°		
26.82	1.13977	6.13	0.99938
14.46	.05194	1.56	.96860
	70.19°		
29.13	1.17539	4.94	1.0826
21.95	.12359	2.50	0.99310
15.83	.07839	2.04	.99026
11.82	.05242	1.25	.98543
6.72	.01938		
	50.04°		
22.01	1.13345	4.94	1.01850
15.68	.08912	2.48	.00343
11.71	.06243	2.02	.00057
6.66	.02962	1.24	0.99575
	25°		
21.82	1.14391	4.93	1.02797
15.55	.09935	2.45	.01271
11.61	.07243	2.00	.00980
6.59	.03927	1.23	.00490
	0°		
21.67	1.18214	4.93	1.03264
15.44	.10664	2.00	.01340
11.55	.07902	1.23	.00832
6.57	.04433		

Herz, 1914

N	d 25°
0	0.9971
1.879	1.0799
2.818	.1201
3.757	.1592
4.174	.1782

Bousfield, 1919

%	d 7°	d 20°	d 33°
20.188	1.1397	1.13427	1.12878
14.821	.10034	.09618	.09111
10.495	.07030	.06625	.06205
7.150	.04752	.04445	.04010
3.667	.02416	.02172	.01777
1.8605	.01219	.01011	.00636
0	0.99993	0.99823	0.99473

Burrows, 1919

%	d	%	d
	25°		
0	0.99707	4.80	1.02858
0.15	.99812	7.50	.04608
0.38	.99978	10.23	.06440
1.00	1.00427	16.67	1.10755
1.93	.01014	20.55	.13458

Pulvermacher, 1920

N	d 25°
0.490	1.0202
0.980	.0419
1.960	.0848
3.920	.1667

Cohen and Moesveld, 1920

%	d
30°	
7.150	1.04114
20.188	.12998

Manchot, Jahrstorfer and Zepter, 1924

c	d	c	d
25°			
5.50	1.0334	19.30	1.1385
8.53	.0540	21.58	.1588
12.09	.0807	23.15	.1734
12.86	.0850		

Goard, 1925

N	d
20°	
1	1.044
2	.088
3	.129
3.8	.152

de Block, 1925

%	d
16°	
0	0.9990
7.13	1.0453
13.7	.0899
23.4	.1586

Cornec and Hering, 1926

%	d	t sat.sol.
26.43	1.179	100°
30.07	.194	75°
33.16	.202	50°
35.70	.211	25°

Hrynakowski, 1927

40.40° d= 1.1945 sat.sol.

Harrison and Perman, 1927

%	d	%	d	%	d
40.0°		50.0°		60.0°	
3.16	1.016	4.23	1.014	3.80	1.006
3.90	.018	8.81	.044	5.15	.014
5.82	.031	10.91	.058	8.08	.032
8.68	.049	14.27	.080	10.91	.053
9.77	.054	21.46	.133	13.17	.067
10.43	.058	26.32	.167	15.92	.086
13.03	.077	30.00	.194	21.61	.127
17.84	.112			24.64	.148
18.80	.118			28.00	.174
22.13	.140			31.30	.197
23.41	.151				
27.30	.178				
29.44	.195				
78.0°		80.0°			
4.14	1.002	9.45	1.031		
5.85	.012	11.47	.041		
7.16	.021	15.13	.068		
8.74	.028	18.59	.094		
10.18	.039	22.73	.124		
12.35	.053	27.24	.157		
13.66	.061	30.01	.177		
17.11	.086	34.00	.215		
21.92	.123				
27.00	.137				
28.38	.148				
32.60	.200				

Scott and Frazier, 1927

26.38% sat.sol. : d^{25}= 1.17786

Flöttmann, 1928

%	t	d
24.73	15	1.1696
25.58	20	.1739
26.46	25	.1780

Herz and Hiebenthal, 1929

N	d 25°	d 70°
0.5	1.0200	1.0016
1.0	.0425	.0236
2.0	.0863	.0662
3.0	.1282	.1074

Perman and Urey, 1929

%	d	%	d	%	d
30°		40°		50°	
2.59	1.011	2.59	1.010	2.61	1.005
2.61	.012	5.37	.026	5.37	.022
5.37	.029	10.40	.059	10.40	.054
10.40	.062	12.23	.072	12.23	.067
12.23	.075	17.91	.110	17.92	.105
17.92	.113	22.19	.141	22.19	.135
22.19	.144				
60°		70°		80°	
2.51	0.998	2.51	0.993	2.51	0.988
5.37	1.016	5.37	1.010	5.37	1.005
11.13	.053	12.23	.054	10.40	.035
12.23	.061	17.92	.094	12.23	.048
17.92	.100	22.19	.124	17.92	.088
22.19	.130			22.19	.118

Geffcken, 1929

m	d	m	d
25°			
0	0.99707	3.0312	1.11953
1.1323	1.04672	4.4104	.16534
1.4927	.06184	4.4246	.16588
2.2073	.08956		

Shibata and Hölemann, 1931

m	d	m	d	m	d
25.0°		35.0°		45.0°	
0	0.99707	0	0.99406	0	0.99024
1.2983	1.05387	1.2982	1.05014	1.2880	1.04584
2.6611	.10707	2.6812	.10290	1.2984	.04585
4.3122	.16245	4.3172	.15803	2.6763	.09813
4.3137	.16249			2.6779	.09823
				4.3185	.15318

Cornec and Krombach, 1932

%	t	d	%	t	d
21.92	0	1.154	33.59	80	1.205
25.57	20	1.174	35.69	100	1.209
28.65	40	1.189	36.50	108.5	1.200
31.29	60	1.199			

Jones (Gr.) and Talley, 1933

N	d	N	d
25°			
0.498450	1.020229	2.011510	1.087400
0.999718	1.042888	2.962076	1.127902

Thomas and Perman, 1934

%	d	%	d
30°			
2.59	1.011	12.23	1.075
5.37	.029	17.92	.113
10.40	.062	22.19	.144

Spacu and Popper, 1934

%	d	%	d
20°			
0	0.99823	16.930	1.11158
24.760	1.16785	14.889	.09720
24.610	.16675	14.850	.09685
24.250	.16404	7.456	.04628

Scott, Obenhaus and Wilson, 1934

%	d
35°	
7.4440	1.104118
17.2876	.10725
26.522	.17419

Gibson, 1935

%	d	%	d
25°			
0.00	0.9979	16.84	1.1082
5.88	1.0328	20.00	.1305
10.91	.0678	23.22	.1540
13.79	.0872	25.52	.1701

Hering, 1936

wt%	mol%	t	d
		sat.sol.	
26.45	7.99	25	1.179
30.05	9.40	50	.194
33.15	10.70	75	.203
35.7	11.83	100	.211

Rodnyanskii and Galinker, 1955

t	d 1N	d 2N	d 3N
25	1.042	1.082	1.119
50	.032	.071	.106
100	.001	.040	.078
150	0.964	.005	.045
200	.924	0.970	.010
250	.878	.930	0.972
300	.826	.879	.925
340	.747	.814	.861

b) Specific volume

Sorby, 1859

t	v/v_0	t	v/v_0
	25%		
0	1.0000	125	1.0623
25	.0087	150	.0802
50	.0196	175	.1000
75	.0322	200	.1228
100	.0464		

Nicol, 1884, 1886 and 1887

mol%	molar volume
	20°
1	1827.70
2.9	1884.58
4.8	1949.04
6.5	2013.36

t	$v/v_{20°}$	t	$v/v_{20°}$
1 mol%		2.9 mol%	
20	100.000	20	100.000
45.4	100.870	45.6	100.958
50.5	101.101	50.3	101.171
56.2	101.374	56.5	101.470
61.3	101.639	61.7	101.737
67.2	101.965	67.4	102.051
72.4	102.271	72.3	102.323
78.5	102.641	78.2	102.671
4.8 mol%		6.5 mol%	
20	100.000	20	100.000
45.6	100.992	45.7	101.036
50.3	101.205	50.6	101.251
56.5	101.513	56.3	101.528
61.6	101.760	61.6	101.782
67.4	102.063	67.3	102.074
72.3	102.327	72.3	102.335
78.2	102.664	78.5	102.670

Jones (Gr.), Taylor and Vogel, 1948

m	$\tau.10^4$ 25°	35°	45°	55°	65°	75°
0.06022	2.6018	3.4705	4.2258	4.9018	-	-
.06022	.6038	.4704	.2268	.8988	-	-
.10053	.6277	.4842	.2296	.8985	5.5122	6.0850
.10053	.6273	.4851	.2285	.8984	.5108	.0837
.10064	.6265	.4881	.2297	.8974	.5115	.0843
1.0294	3.0798	.7496	.3281	.8583	.3513	5.8220
.0308	.1049	.7513	.3292	.8592	.3499	.8240
.0308	.1063	.7499	.3309	.8600	.3516	.8210
3.2731	.6642	4.0365	.3885	.7238	.0436	.3579

Zepernick and Tammann, 1895

t	v/v_0 3.943%	8.055%	15.761%
0	1	1	1
110	1.0536	1.0528	1.0528
120	.0613	.0612	.0605
130	.0701	.0690	.0683
140	.0792	.0780	.0766

Kurochkin, 1929

c	Dv(%)
4	1.4
10	3.6
20	7.4

Pohl, 1852

%	π	%	π
		13.9°	
0	1.000	18.24	0.780
6.32	0.875	23.90	0.687
11.80	0.821		

Schumann, 1887

%	π	%	π
		15°	
2.52	48.1	10.68	40.0
5.35	42.4	16.81	35.4

Röntgen and Schneider, 1888

%	π	%	π
		18°	
0.00	1.0000	9.82	0.8515
0.871	0.9866	15.03	0.7840
2.48	0.9603	19.94	0.7263
4.88	0.9212	25.52	0.6657
7.51	0.8831		

Drecker, 1888

%	π uncorrected	%	π uncorrected
	0°		
2.49	42.6	16.85	34.1
4.40	41.2	19.97	31.7
8.28	38.9	24.31	30.1
13.02	35.4		

Gilbault, 1897

%	π	%	π
	20°		
0	44.37	15	36.24
5	41.49	20	33.90
10	38.76	23.97	32.08

Perman and Urey, 1929

%	π 0-100 atm.	π 100-200 atm.
	30°	
2.59	42.9	41.2
2.61	42.8	41.1
5.37	41.2	39.9
10.40	38.5	37.4
12.23	37.5	36.4
17.92	34.7	33.3
22.19	32.8	31.8
	40°	
2.59	42.4	40.9
5.37	40.8	39.5
10.40	38.4	37.2
12.23	37.1	36.3
17.91	34.4	33.2
22.19	32.6	31.6
	50°	
2.61	42.5	41.0
5.37	40.8	39.8
10.40	38.5	37.3
12.23	37.3	36.4
17.92	34.6	33.5
22.19	32.7	31.9
	60°	
2.51	42.6	41.2
5.37	41.0	40.0
11.13	37.8	37.0
12.23	37.4	36.6
17.92	34.7	33.8
22.19	33.2	32.2
	70°	
2.51	43.1	41.9
5.37	41.4	40.4
12.23	37.7	37.6
17.92	35.4	34.0
22.19	33.5	32.6
	80°	
2.51	44.0	42.8
5.37	42.2	[illegible]
10.40	39.7	38.5
12.23	38.6	37.4
17.92	35.8	34.5
22.19	33.9	32.9

Freyer, 1931

%	π	t	π
20°		20%	
1	45.32	15	33.91
6	41.76	20	33.55
10	39.17	25	33.29
16	35.67	30	33.03
20	33.55	40	32.68
24	31.55	50	32.51

Scott, Obenhaus and Wilson, 1934

%	π
35°	
7.4440	37.83
17.2876	33.35
26.522	29.64

Thomas and Perman, 1934

%	π	%	π
	30°		
2.59	42.8	12.23	37.5
5.37	41.2	17.92	34.7
10.40	38.5	22.19	32.8

Gibson, 1935

%	π (1 - 1000 bars)
25°	
0.00	39.35
5.88	36.32
10.91	34.22
13.79	33.02
16.84	31.67
20.00	30.39
23.22	29.29
25.52	28.46

b) viscosity and surface tension

Grotrian, 1876

%	η 6°	η 15.62°	η 18°	$\tau.10^4$
9.93	1856	1426	1318	225
20.95	1818	-	1347	206

Brückner, 1891

mol%	η 15°	η 20°
0	1143.9	1008.6
0.5	1124.7	999.0
1	1114.0	995.2
1.5	1106.5	992.3
2	1108.5	998.0
2.5	1110.5	1004.5
3	1121.2	1017.5
3.5	1134.6	1031.4

Taylor and Ranken, 1903-04

molarity	η (water^0=1) 0°	15°	25°
1	0.931	0.622	0.502
2	.886	.615	.507
3	.880	.625	.517

Bousfield, 1905

%	η 15°	η 18°	η 21°
21.0000	1122.4	1056.7	998.5
18.0047	1114.4	1047.0	987.1
15.0095	1107.5	1039.7	977.1
12.9993	1102.5	1033.3	971.0
12.0010	1106.4	1034.8	971.5
11.4589	1102.9	1032.8	969.8
10.0001	1106.7	1033.6	968.9
9.0026	1108.0	1035.7	970.7
5.9998	1114.5	1039.3	970.4
4.00016	1121.1	1041.9	971.2
3.0003	1124.4	1044.5	973.1
2.0000	1130.0	1046.5	973.8
1.00005	1134.5	1050.4	977.2
0.0000	1135.6	1051.4	976.9

Getman, 1907

%	η	%	η
	18°		
4.91	1053	17.61	1089
9.44	1056	21.77	1106
13.68	1079	25.83	1120

c	η 10°	η 30°	η 50°
0	1310	800	550
11.28	1250	820	590
27.14	1250	860	650

Lorenz, 1912

N	t	η
20.91	790	1420
20.38	835	1210
19.97	920	990
19.14	1035	710

Herz, 1914

N	η (water=1)	N	η (water=1)
	25°		
0	1	3.757	1.067
1.879	1.0799	4.174	1.097
2.818	1.1201		

Pulvermacher, 1920

N	η (water=1)	N	η (water=1)
	25°		
0.490	0.9962	1.960	1.004
0.980	0.9917	3.920	.071

Hrynakowski, 1927

40.40° η= 809 sat.sol.

Tammann and Rabe, 1928

%	η 0°	10°	30°	75°
6.73	1655.3	1258.8	806.1	404.1
13.02	1609.9	1239.2	822.9	427.3
18.15	1569.9	1243.0	837.4	447.6
23.00	-	1267.4	870.6	475.4
27.03	-	-	-	501.1

Herz and Hiebenthal, 1929

N	η 25°	70°
0.5	893.2	424.5
1.0	889.4	436.5
2.0	899.4	458.1
3.0	919.3	490.6

Jones (Gr.) and Talley, 1933

N	η (water=1)
25°	
0.498450	0.99760
.999718	.99702
2.011510	1.00773
.962076	.03211

Jacopetti, 1940

t	η	t	η
6.642%		12.976%	
25.00	892.4	24.70	907.2
30.40	801.9	30.35	811.8
50.10	568.2	50.13	589.0
73.50	409.9	75.88	432.5
18.256%		22.944%	
20.64	997.0	21.05	1012.5
30.37	833.8	30.38	860.3
50.10	614.4	50.11	641.6
72.30	465.0	73.22	488.5

Clack, 1908

%	k^{*}
0°	
10.08	95.4
19.89	97.2

k = velocity of diffusion

Volkmann, 1882

%	σ
15-16°	
7.19	75.4
15.21	77.1
24.66	75.4

Rother, 1884

%	σ	%	σ
15°			
2.50	72.73	10.01	74.43
4.95	73.21	14.84	75.61
7.45	73.79	19.62	76.98
		23.04	78.07

Traube, 1885

c	%	σ
15°		
10	9.41	74.24
20	17.85	78.55

Ochsé, 1890

t	a^2	t	a^2	t	a^2
0%		5g/100cc *		10g/100cc	
0	80.65	6	62.23	4	59.68
4	77.52	15	60.76	15	58.31
8	76.36	25	59.09	25	56.94
15	72.91	35	57.23	35	55.47
25	70.17	45	55.76	45	54.00
40	64.19	55	54.59	55	53.12
55	60.27				
15g/100cc		20g/100cc			
4	58.42	5	52.23		
9	55.17	10	52.14		
15	54.19	15	51.35		
26	53.61	25	50.86		
40	52.82	40	50.47		
55	51.94	55	49.98		

g/100cc at each t .

Sentis, 1897

mol%	t	σ	mol%	t	σ
6	21.2	77.8	1	16.45	74.5
5	21.8	76.7	1	12.2	75.3
5	16.4	77.8	1	19.4	73.8
3	16.3	76.0	0	13.5	74.0
2	16.55	75.3	0	25.1	72.3
2	12.3	75.8			

Linebarger, 1899

%	σ
	25°
9.09	72.05
10.71	72.33
16.67	73.61
23.08	75.39
26.44	76.97

Grubowsky, 1904

%	σ 10°	σ 30°
0	74.05	71.06
9.47	75.90	72.97
18.27	78.07	75.06
23.32	79.56	76.68

Forch, 1905

N	t	σ	N	t	σ
0.606	15.4	77.72	1.716	17.0	79.05
1.010	15.4	78.29	2.12	17.5	79.49
1.430	16.8	78.63	3.00	16.9	80.88

Stocker, 1920

%	σ
	18°
0	72.56
6.73	74.16
12.65	75.11
21.06	77.67

de Block, 1925

%	σ
	16°
0	73.11
7.13	74.50
13.7	75.89
23.4	78.37

Goard, 1925

N	σ
	20°
1	74.55
2	76.13
3	77.87
3.8	79.17

Hrynakowski, 1927

40.40° σ= 74.78 sat.sol.

Optical and electro-magnetic properties .

Kremers, 1857

%	n_D
17°	
0	1.3332
23.7	1.3646

Fouqué, 1867

%	D n_D (.10^5)
9-93°	
0.7	16
7.15	16
19.65	17

Börner, 1869

t	$n_{H\alpha}$	t	$n_{H\beta}$	t	$n_{H\gamma}$
10%					
44.1	1.340265	43.1	1.346598	42.3	1.350210
39.4	.340913	38.55	.347344	37.8	.350926
35.4	.341542	34.5	.347989	34.1	.351426
30.5	.342208	30.8	.348526	31.3	.351820
27.46	.342658	27.5	.348992	27.5	.352427
20%					
44.8	1.350322	44.1	1.357103	43.3	1.360947
40.3	.351069	40.2	.357744	39.7	.361516
35.25	.351820	35.7	.358452	36.3	.362138
31.9	.352356	31.5	.359133	31.2	.362866
28.85	.352767	28.75	.359560	28.7	.363221
30%					
44.1	1.359424	43.1	1.366520	42.1	1.370553
38.8	.360271	39.1	.367228	39.9	.371012
33.0	.361125	33.3	.368184	34.3	.371895
28.9	.361801	28.9	.368785	27.4	.372883
24.1	.362440	24.3	.369457	24.7	.373307
21.25	.362937	21.3	.369987	21.5	.373825

Le Blanc, 1889

%	n_D
20°	
0	1.33325
5.05	.34013
20.55	.36173

Bender, 1890

N(15°)	t	$n_{H\alpha}$	t	n_D
0.0	20	1.33100	18.8	1.33297
0.5	16	.33644	18	.33644
1.0	18.5	.34079	17.8	.34079
1.5	16.5	.34551	17.2	.34551
2.0	17.5	.34990	17	.34990
2.5	18.5	.35422	20.6	.35422
3.0	20.3	.35794	19.8	.35794
3.5	21.3	.36134	20.9	.36134

N(15°)	t	$n_{H\beta}$	t	$n_{H\gamma}$
0.0	20.2	1.33703	20.3	1.34021
0.5	15.8	.34245	16.9	.34570
1.0	18.6	.34709	18.0	.35049
1.5	17.3	.35191	17.1	.35547
2.0	17.7	.35650	17.6	.36000
2.5	17.2	.36087	18.7	.36469
3.0	20.1	.36478	19.8	.36861
3.5	21.0	.36829	20.8	.37209

Doumer, 1892

%	n_D
15°	
6.29	1.3421
11.70	.3497
16.70	.3566
23.08	.3652
30.47	.3759

Borgesius, 1895

%	n_D	%	n_D
20°			
0	1.333000	4.44	1.333640
0.58	.333078	8.50	.334266
1.15	.333159	15.58	.335515
2.28	.333321		

Bender, 1900

t	n H_α	n H_β	n H_γ
		0.5 N at 15°	
10	1.33683	1.34294	1.34632
15	.33644	.34257	.34590
20	.33598	.34213	.34543
25	.33550	.34157	.34494
30	.33498	.34102	.34438
35	.33436	.34055	.34377
40	.33367	.33975	.34309
		1 N	
10	1.34162	1.34793	1.35137
15	.34117	.34745	.35090
20	.34071	.34697	.35039
25	.34017	.34643	.34983
30	.33956	.34578	.34925
35	.33890	.34516	.34857
40	.33821	.34441	.34781
45	.33739	.34357	.34697
50	.33652	.34273	.34613
55	.33582	.34192	.34525
60	.33495	.34109	.34439
65	.33408	.34022	.34351
70	.33335	.33935	.34263
		2 N	
10	1.35069	1.35722	1.36093
15	.35015	.35667	.36035
20	.34955	.35612	.35978
25	.34896	.35550	.35912
30	.34840	.35495	.35848
35	.34775	.35425	.35784
40	.34705	.35358	.35719
45	.34629	.35279	.35651
50	.34553	.35200	.35568
55	.34468	.35117	.35484
60	.34390	.35035	.35391
65	.34307	.34943	.35304
70	.34212	.34841	.35206
		3 N	
10	1.35919	1.36605	1.36985
15	.35861	.36541	.36923
20	.35794	.36479	.36863
25	.35728	.36416	.36794
30	.35668	.36352	.36730
35	.35606	.36289	.36664
40	.35537	.36216	.36589
45	.35464	.36148	-
50	.35380	.36063	.36442
55	.35302	.35978	.36357
60	.35218	.35993	.36273
65	.35145	.35812	.36191
70	.36062	.35728	.36104

Wagner, 1903

c	n_D	c	n_D
	17.5°		
0	1.33320	12.571	1.34910
0.286	.33358	.879	.34947
.573	.33397	13.187	.34984
.858	.33435	.496	.35021
1.144	.33474	.807	.35058
.430	.33513	14.118	.35095
.720	.33551	.429	.35132
2.009	.33590	.740	.35169
.300	.33628	15.051	.35205
.591	.33667	.362	.35242
.882	.33705	.673	.35279
3.173	.33743	.985	.35316
.464	.33743	16.297	.35352
.755	.33781	.609	.35388
4.046	.33820	17.239	.35425
.338	.33858	.554	.35461
.635	.33896	.869	.35497
.940	.33934	18.184	.35533
5.241	.33972	.499	.35569
.542	.34010	.814	.35606
.843	.34048	19.130	.35642
6.144	.34086	.446	.35678
.445	.34124	.762	.35714
.746	.34162	20.078	.35750
7.047	.34199	.394	.35786
.349	.34237	.710	.35822
.655	.34275	21.026	.35858
.961	.34313	.342	.35894
8.267	.34350	.658	.35930
.573	.34388	.974	.35966
.880	.34426	22.290	.36002
9.187	.34463	.606	.36038
.494	.34500	.922	.36074
.801	.34537	23.232	.36109
10.108	.34575	.542	.36145
.415	.34612	.853	.36181
.720	.34650	24.164	.36217
11.031	.34687	.475	.36252
.339	.34724	.786	.36287
.647	.34761	25.097	.36323
.955	.34798	.408	.36359
12.263	.34836	.719	.36394
	.34873	16.924	.36429

Dinkhauser, 1905

%	n_D
18°	
2.290	1.33560
5.843	.34162
8.717	.34554
12.481	.35078
18.521	.35885

Zecchini, 1905

%	t	n_D
0	8.0	1.33376
5.805	8.8	.34162
12.422	7.7	.35096
15.359	6.9	.35556
17.628	7.4	.35838
19.769	8.6	.36165

Chéneveau, 1907

%	n_D	%	n_D
	15°		
0	1.3334	15.42	1.3545
2.77	.3371	17.74	.3578
5.49	.3408	19.97	.3610
8.10	.3443	22.14	.3641
10.61	.3477	24.26	.3670
13.06	.3512		

%	n C	n D	n Tl	n F	n G
23.57	1.36280	1.36594	1.36818	1.37079	1.37482
13.99	.35010	.35209	.35423	.35662	.36026
4.97	.33778	.33966	.34165	.34393	.34726

Getman and Wilson, 1908

%	n_D	%	n_D
	20°		
0	1.33298	15	1.35368
5	.33975	20	.36127
10	.34650	24	.36960

Guerdikova, 1910

%	n_D	%	n_D
	25°		
0	1.33255	19.143	.3587
10.345	.3466	25.830	.3683

Heydweiller, 1909 and Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34299
0.1	.33423	2.0	.35215
0.2	.33525	3.0	.36053
0.5	.33825		

Baxter, Boylston and al., 1911

%	n_D	%	n_D
	25°		
0	1.33327	7.1573	1.34215
0.9466	1.33376	7.1699	.34215
.9716	.33380	9.0221	.34463
.9769	.33380	.4153	.34517
.9983	.33382	.5004	.34528
1.0013	.33388	13.758	.35122
2.4297	.33577	.755	.35122
.4623	.33581	17.852	.35678
4.6474	.33871	.898	.35687
.8150	.33896	21.522	.36202
.8318	.33896	.770	.36238
.8340	.33896		

Shippy and Burrows, 1918

%	n_D
	25°
5	1.33931
10	.34643
15	.35335
20	.35992

Barnes and Johnson, 1902

%	κ	%	κ
25.5°			
	24.83	11.87	121.5
0.97	15.38	14.20	147.2
1.61	33.20	18.81	189.5
3.10	55.33	21.92	222.5
5.50	78.28	23.91	239.7
7.72	96.99	29.64	291.5
9.50			

Pulvermacher, 1920

N	n_D
25°	
0.490	1.3382
0.980	.3429
1.960	.3513
3.920	.3674

Flöttmann, 1928

%	t	n_D
24.73	15	1.36797
25.58	20	.36853
26.46	25	.36901

Geffcken, 1929

N	n_{He}	N	n_{He}
25°			
0	1.332590	3.0312	1.357695
1.1323	.342940	4.4104	.366705
1.4927	.346121	4.4246	.366785
2.2073	.351735		

Shibata and Höleman, 1931

N	n_{He}	N	n_{He}	N	n_{He}
25°		35°		45°	
0	1.33270	0	1.33149	0	1.33000
1.2983	.34460	1.2982	.34325	1.2980	.34168
2.6611	.35533	2.6812	.35393	1.2984	.34166
4.3122	.36623	4.3172	.36479	2.6779	.35232
4.3137	.36626			2.6763	.35230
				4.3185	.36317

Spacu and Popper, 1934

%	n_{He}	%	n_{He}
20°			
24.760	.3668197	14.889	.3529790
24.610	.3666130	14.850	.3529040
24.250	.3661130	7.456	.3426270
16.930	1.3558640	0	1.3224865

Akhumov and Golovkov, 1935

%	n_D	%	n_D
15°			
0	1.3334	13.60	.3548
3.50	.3405	19.80	1.3629
7.10	.3432	24.76	.3706

Oppenheimer, 1898

%	$(\alpha)_{magn.}$
20°	
6.64	1.32
12.82	1.33
17.66	1.34
23.93	1.31

Agerer, 1908

%	$(\alpha)_{magn.}$
18°	
7.505	1.337
11.958	1.330
17.098	1.329

Guerdikova, 1910

%	$\alpha^D_{magn.}$	%	$\alpha^D_{magn.}$
25°			
0	5°068	19.143	5°834
10.345	5°368	25.830	6°175

Okazaki, 1933

%	Verdet's constant (3441 Å)
28°	
4.96	0.04653
10.05	.04905
14.98	.05145
17.11	.05275
19.98	.05458
24.98	.05744

Okazaki, 1936

%	Verdet's constant (D)
25°	
0	0.013075
5.459	.01380
8.495	.01421
12.89	.01481
16.93	.01539
20.74	.01592
23.87	.01641

Okazaki, 1951

%	Verdet's constant (D)
35°	
0	0.013047
11.12	.01437
14.11	.01480
17.33	.01523
21.33	.01577
25.11	.01631

Electrical properties

Grotrian, 1876

%	ϰ	τ.10⁴
	18°	
9.93	1341	186
20.95	2889	165

Kohlrausch and Grotrian, 1875

%	ϰ 0°	ϰ 18°
5	453	686
10	920	1350
15	1395	2008
21	1991	2724

Kohlrausch, 1879

%	ϰ	Dϰ/ϰ18°
18°		
5	687	0.0202
10	1351	.0189
15	2008	.0180
20	2662	.0169
25	2793	.0167

Bouty, 1887

N (15°)	ϰ	τ.10⁴
	0°	0-30°
0.001	0.7876	333
.01	7.546	333
.1	70.93	327
.2	138.5	326
.5	328.0	302
1.0	648.7	291
2.0	1285	259
3.0	1933	230

Wood, 1895

N	λ	N	λ
		0°	
0.001	76.35	0.06	68.70
.002	75.90	.12	66.47
.004	75.30	.25	64.46
.008	73.90	.5	62.49
.015	72.60	1	61.10
.03	70.66		

Dennhardt, 1899

N	λ 0°	λ 10°	λ 18°
0.1	71.4	92.4	112.0
0.5	68.0	87.2	101.0
1	65.5	83.2	98.2
2	62.9	79.8	92.5
3	60.9	76.2	88.7

Barnes and Johnson, 1902

%	κ	%	κ
	25.5°		
29.64	291.5	9.50	96.99
23.91	239.7	7.72	78.28
21.92	222.5	5.50	55.33
18.81	189.5	3.10	33.20
14.20	147.2	1.61	15.38
11.87	121.5	0.97	24.83

Taylor and Ranken, 1903 - 1904

N	λ
0°	
1	65.4
2	63.1
3	62.4

Jones, 1904

N	λ	N	λ
	0°		
0.05	72.37	0.50	65.21
.10	69.53	1.00	61.07
.20	68.71	2.00	60.98

Inclan, 1906

t	λ	t	λ
7.78 mol%		4.05 mol%	
3	35.31	2	66.03
11.3	69.58	8	77.31
19.5	80.08	19	95.43
33	97.77	33.3	121.62
40.8	109.81	40.5	133.66
50	120.02	50.2	150.16
58.2	129.76	61	165.03
70	141.28	69	178.63
78.3	149.84	78.5	194.04
89	158.81	89.3	209.14
94	167.81		

t	λ	t	λ
2.74 mol%		2.07 mol%	
4	64.84	2.5	66.785
9.3	82.40	10.3	79.134
20.2	99.69	19.4	96.134
32.9	120.90	33.2	122.59
41.8	136.47	41.5	142.95
49.3	145.17	50.3	154.74
59	162.14	60.1	165.32
68.6	180.46	68.5	188.50
78.3	193.07	78.6	205.22
87.1	212.90	87	221.54
94.5	221.99	94	239.20

Jaquerod, 1909

t	κ	t	κ
	70.4g/100 cc		
25.3	1062	80.4	2134
54.8	1635		
	140.8g/100 cc		
25.2	2001	81.4	3888
55.0	2997		
	211.1g/100 cc		
26.2	2888	79.4	5260
54.7	4160		
	281.6g/100 cc		
25.6	3627	80.0	6513
53.7	5127		

Clausen, 1912

N	τ	λ
	30°	
0.498	0.0637	127.8
1.005	.1218	121.2
2.013	.2274	113.0
3.004	.3192	106.3

Lorenz, 1912

N	t	λ
20.91	790	103.29
20.38	835	108.84
19.97	920	121.18
19.14	1035	138.97

Jacopetti, 1940

t	κ	t	κ
6.642%		12.976%	
25.00	1045	24.70	1937
30.40	1138	30.35	2169
50.10	1518	50.13	2828
73.50	1923	73.88	3565
18.256%		22.944%	
20.64	2619	21.05	3207
30.37	2999	30.38	3688
50.10	3826	50.11	4716
72.30	4841	73.22	5841

Rodnyanskii and Galinker, 1955

t	λ		
	1 N	2 N	3 N
25	111.6	105.7	101.0
50	176	157	143
100	286	245	215
150	378	319	277
200	449	378	323
250	490	411	348
300	500	419	353
340	473	392	328

Chambers, Stokes and Stokes, 1956

N	λ	N	λ
1st method		2nd method	
25°			
0.14075	126.64	0.12004	127.67
.19141	124.43	.14968	126.13
.28307	121.51	.18588	124.58
.36417	119.71	.24678	122.53
1.5329	108.08	.32330	120.54
2.0000	105.29	.40698	118.81
2.6082	101.72	.53211	116.79
3.0645	99.08	.66460	115.10
3.6951	95.37	.76142	114.03
4.0000	93.46	.93485	112.41
		1.1934	110.35
		.5642	107.87
		.9155	105.74
		2.3789	103.07
		.6188	101.70
		3.4686	96.72
		.7046	95.29

Thermal constants

Specific heat.

Schüller, 1869

%	U	%	U
20°			
3.85	0.9558	16.67	0.8195
7.41	.9140	19.35	.7935
10.71	.8876	21.88	.7680
13.79	.8503	24.24	.7476

Winkelmann, 1873

%	U	%	U
at room temp.			
3.04	0.9625	15.6	0.8448
4.22	.9500	20.2	.8078
5.58	.9341	25.2	.7760
8.77	.9041	29.4	.7529
11.60	.8773		

Thomsen, 1886

mol %	U	mol %	U
18°			
0	1	1.9	0.904
0.5	0.970	3.2	0.850
1.0	0.948	6.3	0.761

Jaquerod, 1909

g/1	U	g/1	U
at room t.			
70.4	0.9157	211.2	0.7831
140.8	0.8399	281.6	0.7318

W.Q and C.E. Bousfield, 1919

%	U		
	7°	20°	33°
0.000	1.008	1.004	1.004
1.8605	0.981	0.977	0.978
3.667	.955	.954	.956
7.150	.910	.911	.914
10.495	.874	.877	.891
14.821	.828	.832	.837
20.188	.770	.775	.780

Cohen and Moesveld, 1920

%	U	%	U
19°			
0	0.9992	20.188	0.7686
7.150	0.9077		

Randall and Rossini, 1929

m	U	m	U
25°			
0	0.9979	5.0	0.9518
0.1	.9969	7.5	.9312
0.2	.9959	10.0	.9118
0.5	.9929	12.5	.8936
1.0	.9880	15.0	.8765
2.0	.9785	20.0	.8441
3.5	.9649	25.0	.8168

Barnes and Maas, 1930

19.80%

− 78.5° and + 25°

t initial	t final	U*
0.0	+ 25	19.086
0.0	"	19.073
− 8.6	"	25.841
− 9.0	"	26.331
−14.7	"	100.93
−42.9	"	112.30
−44.0	"	112.79
−78.5	"	124.02
−78.5	"	124.15

* in cal/g for (t fin − t init.)

Clews, 1936

t	U	t	U	t	U
2.5 N		2 N		1.5 N	
15.61	0.8138	14.72	0.8390	14.94	0.8722
19.14	.8147	17.96	.8405	18.95	.8735
23.32	.8169	21.13	.8422	24.96	.8750
25.79	.8177	26.22	.8437	28.57	.8755
31.66	.8201	30.47	.8452	32.90	.8765
37.20	.8223	33.81	.8460	38.24	.8773
40.53	.8242	36.65	.8471	45.06	.8794
		40.04	.8490		
0.75 N		0.5 N		0.1 N	
21.61	0.9318	16.11	0.9556	16.37	0.9885
28.94	.9324	20.95	.9557	19.30	.9871
34.48	.9333	26.76	.9556	23.86	.9872
40.85	.9341	31.48	.9558	29.11	.9871
44.41	.9342	33.90	.9554	33.73	.9875
50.64	.9350	38.42	.9557	38.82	.9870
		42.78	.9558	43.15	.9882
		16.27	.9560		
		22.25	.9557		
		29.81	.9558		
		36.34	.9554		
		42.63	.9561		

Nikolaev, Kogan and Ogorodnikov, 1936

%	U	%	U
25°			
0.50	0.9908	15.08	0.9266
4.00	.9468	22.68	.7465
6.08	.9228	24.52	.7274

Rutskov, 1948

%	U 25°	U 50°	U 75°
0.740	0.9880	0.9885	0.9912
3.594	-	.9532	.9565
9.46	0.8829	.8857	.8890
15.71	-	.8190	.8211
25.64	0.7193	.7203	.7208

Kapustinskii, Yakuchevskii and Drakin, 1953

%	U	%	U
0	0.9980	7.97	0.8997
3.76	0.9498	11.31	0.8635
5.92	0.9237		

Heat of dissolution .

Winkelmann, 1873

%	Q diss (by 1 g KCl) 0°	50°
3.04	-67.01	-44.7
4.22	66.48	46.1
5.59	65.32	44.4
8.77	62.81	44.7
11.60	61.47	43.9
15.60	60.21	44.1
20.20	58.95	43.2
25.20	57.14	42.9
29.40	-55.31	-42.3

Scholz, 1892

N	Q diss.(cal/g)
0.0625	-71.59
.125	71.29
.25	70.84
.5	69.59
1.0	66.95
2.0	63.18
3.6	58.29
8.0	35.31

C..son, 1901

t	Dt 27.8g/100cc + 400cc H_2O
13.4	-0.465
29	-0.31
40.3	-0.21
41.5	-0.20
46.3	-0.145
56.3	-0.07
92.8	+0.18

Vareli-Thevenet, 1902

%	Q diss.(cal/g KCl) 1st series	2nd series
0.2	-79.0	-
0.4	77.8	-
0.6	77.4	-
0.8	75.7	-
1	74.0	-
2	72.6	-75.0
4	69.0	69.7
6	67.6	68.4
8	66.8	66.9
10	64.3	65.6
12	60.3	63.8
14	59.1	62.0
16	56.1	59.7
18	52.7	57.3
20	50.6	55.1
28	45.9	50.1
29	-20.1	-
30	-	-12.1

Bishop, 1908

m initial	m final	Q dil. (by mole KCl)	Q dil integral (by mole KCl)
4.492	3.804	48.0	48.0
4.492	3.808	48.7	48.7
3.804	3.127	55.8	103.8
3.127	2.564	47.1	150.9
2.564	2.053	46.1	197.0
2.053	1.670	44.1	241.1
4.492	0.673	372.6	372.6
4.492	0.666	384.8	384.8
4.492	0.718	395.7	395.7
1.461	0.058	60.0	306
1.050	0.036	53.0	358
1.461	0.062	73.0	418

Stearn and Smith, 1920

initial molality	g sol.	moles aq. added	Q dil. (by mole aq.)
3.16	8.410	39.0	-20.03
		26.55	22.76
		19.4	24.94
		14.4	26.48
		9.4	27.79
2.6	8.000	40.0	7.24
		30.0	8.22
		20.0	8.87
0.8	8.000	40.0	10.33
		30.0	1.77
		20.0	2.44
		10.0	2.78
0.4	8.000	34.95	-0.43
		30.0	0.45
		25.0	0.564
		15.0	0.67
0.2	7.850	40.0	0.069
		38.2	0.076
		34.0	0.090
		30.0	0.103
		30.0	0.108

Mondain-Monval, 1923

	0°	18°
Q diss(0.85mol)	-5010	-4270
Q dil(sat.sol.- by mole KCl)	-950	-590
Q dil(1 mole H_2O added)	-40	-38
Q dil(1 mole KCl added)	-570	-465

Lehtonen, 1922

N	Q diss.(cal/g)	N	Q diss.(cal/g)
	0°		
0.0625	-71.397	1.0000	-66.801
.1250	-71.114	2.0000	-63.016
.2500	-70.726	4.0000	-57.038
.5000	-69.409		

Next Page

Wüst and Lange, 1923 and 1924

N	Q diss. integral	N	Q diss. integral
	25°		
0.6642	-4231	0.331	-4231
1.3228	4177	0.993	4123
1.970	4123	1.64	4010
2.607	4069	2.29	3906
3.242	4014	2.93	3794
3.863	3963	3.55	3715
4.479	3914	4.18	3621
5.093	3868	4.79	3548
5.693	3826	5.39	3498
6.294	3786	5.99	3430
6.882	3749	6.57	3385
7.450	-3715	7.16	-3355

Harrison and Perman, 1927

%	Q dil *	%	Q dil*
	40.0°		
100	+18.92	22.46	0.88
27.26	1.51	19.54	0.62
25.41	1.24	18.24	0.58
23.86	1.05	17.00	0.47
	50.0°		
100	+18.49	16.31	0.31
27.86	1.18	12.33	0.12
25.79	1.10	9.17	0.04
19.03	0.53		
	60.0°		
100	+18.28	18.53	0.30
28.85	0.56	16.08	0.18
25.26	0.47	12.64	0.06
22.16	0.40		
	70.0°		
100	+18.42	22.10	+0.07
30.87	-0.23	21.43	+0.06
29.67	-0.14	18.86	+0.04
28.86	-0.08	18.28	+0.02
25.71	-0.01	16.92	+0.02
22.97	+0.03	16.29	+0.01
	80.0°		
100	+17.84	20.32	-0.16
32.28	-0.99	17.66	-0.06
30.57	-0.77	14.87	-0.03
26.77	-0.54	12.7	-0.01
23.60	-0.22		

* Q dil = heat evolved on addition of 1 g water to an infinite quantity of solution.

Kaganovich and Miscenko, 1951

m	Q diss. (by mole KCl)	
50°		
0.138	-3420	
.274	3408	
.415	3393	
.467	3383	
.521	3380	
.826	3353	
1.110	3351	
.666	3310	
2.216	3256	
.828	3249	
3.325	3203	
.877	3226	sic
4.436	3210	
.744	-3212	

Kapustinskii and Ruzavin, 1955

%	c	%	c
	25°		
4.92	1433	18.1	1375
9.43	1412	24.6	1335
13.3	1394		

heat conductivity coefficient c $(.10^6)$

Water (H_2O) + Potassium bromide (KBr)

Heterogeneous equilibria

Tammann, 1885

t	p 0%	19.76%	26.15%	30.77%	40.68%
66.07	196.8	184.9	178.6	173.0	161.6
70.88	242.8	227.7	220.1	213.0	199.4
75.71	298.0	278.5	269.3	261.1	243.4
81.71	380.8	356.0	343.6	332.7	310.0
88.78	502.3	469.5	453.2	438.0	406.9
91.61	559.0	522.8	503.7	488.0	452.8
95.22	639.2	598.8	576.5	558.5	517.5
100.34	769.4	718.1	695.0	671.0	-

Tammann, 1885

%	p	%	p
100°			
5.73	748.8	28.99	671.1
9.30	739.2	34.66	640.4
16.81	718.5	38.00	623.0
24.64	690.2	41.23	602.9
26.85	680.1	46.40	567.5

Dieterici, 1891

%	p
0°	
0	4.620
10.63	.476
19.22	.324
26.31	.186
32.25	.037

Lannung, 1934

m	p	m	p
18°			
0.9977	15.27	1.803	14.60
0.5481	15.20	2.666	14.17
1.052	14.97	3.289	13.84
.124	14.93	4.014	13.47
.245	14.87	4.619	13.11
.477	14.77	5.041	12.79
.682	14.67	sat.sol.	12.70

Baroni, 1893

%	b.t.	%	b.t.
1.23	100.101	11.30	101.017
2.42	100.191	15.61	101.490
3.59	100.288	19.86	102.063
5.02	100.413	24.08	102.689
6.94	100.584	27.40	103.265
9.05	100.781		

Buchanan, 1899

m	b.t.	m	b.t.
0.43	100.01	1.30	100.81
0.56	100.11	1.51	101.02
0.65	100.21	1.71	101.22
0.77	100.31	2.07	101.62
0.89	100.41	2.54	102.13
0.98	100.51	3.38	103.14
1.10	100.61	4.55	104.66

Kahlenberg, 1901

%	b.t.	%	b.t.
760 mm			
2.54	100.21	3.27	100.25
5.21	100.44	5.36	100.46
8.76	100.78	8.61	100.74
14.16	101.38	11.88	101.09
18.96	102.01	15.23	101.47
22.24	102.46	18.24	101.90
24.96	102.95	20.94	102.30
27.80	103.50	23.34	102.74
30.27	104.01	25.79	103.19
32.00	104.41	28.46	103.74
33.86	104.88	30.85	104.36
		32.92	104.84

Jablezynski and Kon, 1923

N	b.t.	N	b.t
0.3504	100.326	1.0406	100.995
.5239	100.492	.1987	101.152
.6990	100.659	.3749	101.331
.8708	100.827		

Johnson and Smith, 1941

m	D b.t. 60°	70°	80°	90°	100°
0.1	+0.0724	+0.0774	+0.0828	+0.0884	+0.0944
0.2	.1430	.1532	.1639	.1751	.1868
0.3	.2137	.2291	.2452	.2619	.2794
0.4	.2848	.3054	.3267	.3491	.3723
0.5	.3563	.3820	.4088	.4367	.4657
0.6	.4284	.4594	.4915	.5250	.5598
0.8	.5741	.6156	.6587	.7038	.7507
1.0	.7220	.7746	.8293	.8861	.9453
1.5	1.104	1.185	1.269	1.356	1.447
2.0	.501	.612	.727	.846	.970
2.5	.914	2.057	2.204	2.357	2.516
3.0	2.341	.517	.698	.886	3.081
3.5	.783	.994	3.210	3.435	.668
4.0	3.241	3.487	.741	4.004	4.276
5.0	4.208	4.531	4.862	5.206	5.563

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.2356	0.2347	1.564	1.545
.3252	.3237	.865	1.846
.3587	.3561	2.040	2.016
.5567	.5535	.708	.673
.5651	.5621	3.326	3.292
.5977	.5926	.788	.742
.8002	.7958	4.114	4.069
1.053	1.045	.731	.675
.340	.330	.810	.755
.410	.395		

m_1= molality of potassium chloride

m_2= molality of potassium bromide

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.928	1.4	0.914
0.2	0.916	1.6	0.917
0.3	0.910	1.8	0.922
0.4	0.906	2.0	0.927
0.5	0.904	2.5	0.941
0.6	0.904	3.0	0.955
0.7	0.904	3.5	0.969
0.8	0.905	4.0	0.984
0.9	0.906	4.5	1.000
1.0	0.907	5.0	1.015
1.2	0.910		

Kremers, 1856

%	f.t.	%	f.t.
34.84	0	45.87	60
39.22	20	48.31	80
42.73	40	55.56	100

sat.sol. b.t.=112°

Rüdorff, 1862

%	f.t.
4.53	-1.4
10.04	-3.25
16.54	-5.75
23.73	-9.00

Guthrie, 1876

%	f.t.	%	f.t.
10	-3	34	-5.0
20	-7.1	35.03	0
30	-12.0	39.7	+20
32.15	-13.0 E	43.2	+40
33	-9.8		

De Coppet, 1883

%	f.t.	%	f.t.
31.58	-13.4	42.68	37.9
33.18	-6.2	43.50	43.15
34.78	0.0	43.89	45.45
35.73	3.4	44.60	50.5
36.17	5.2	45.28	54.8
37.90	12.65	46.05	60.15
37.95	13.0	46.87	66.75
38.07	13.3	47.56	71.45
39.06	18.3	47.99	74.85
40.59	26.05	49.31	86.5
41.30	30.0	50.70	97.9
		52.45	110.0

Tilden and Shenstone, 1884

%	f.t.
54.73	140
59.28	181

Etard, 1894

%	f.t.	%	f.t.
31.0	-12	45.5	55
31.5	-10	48.7	77
35.7	+3	54.1	140
41.6	+32	58.5	173
47.7	39	61.6	220

Jones, 1904 and Jones and Getman, 1904

molarity	f.t.	molarity	f.t.
0.099	-0.335	1.598	-5.850
0.199	-0.670	1.998	-7.440
0.399	-1.345	2.397	-9.000
0.799	-2.730	3.196	-12.820
1.199	-4.270		

Meusser, 1905

%	f.t.	%	f.t.
20.0	-6.5	31.2	-11.5
23.8	-8	32.2	-8
26.3	-8.5	33.3	-5
29.5	-10.5	34.5	0
31.0	-11	35.7	3.5
		38.3	10.5

Meyerhoffer, 1905

%	f.t.	%	f.t.
20.0	-6.5	32.2	-8.0
23.8	-8.0	33.3	-5.0
26.3	-8.5	34.5	0.0
29.5	-10.0	35.7	+3.5
31.0	-11.0	38.3	+10.5
31.2	-11.5		

Rodebush, 1918

%	f.t.	%	f.t.
10.33	-3.19	27.71	-10.55
16.96	-5.61	28.32	-10.90
19.23	-6.57	31.34	-12.60 E
23.59	-8.51		

Ceitlin, 1925

N	f.t.
4.28	10.2
4.39	20

Scott and Frazier, 1927

40.57%   f.t.=25°

Scatchard and Prentiss, 1933

m	f.t.	m	f.t.
0.000407	-0.0015	0.17833	-0.6057
.001077	.0039	.19254	.6544
.001824	.0066	.22465	.7610
.003996	.0145	.26581	.8918
.016397	.0587	.40910	1.364
.026573	.0943	.50792	.687
.039203	.1378	.57808	.914
.054501	.1906	.64269	2.125
.081145	.3015	.80717	2.659
.099459	.3433	.97516	3.200
.12646	.4344	1.2086	-3.952
.15791	.5396		

Hering, 1936

wt%	mol%	f.t.
31.27	6.44	-12.5 E
40.65	9.40	25
44.85	10.96	50
48.3	12.39	75
51.2	13.72	100

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
51.5	103	67.7	301
57.5	167	69.4	316
61.1	208	69.0	330
63.1	232	70.5	336
65.2	251	70.7	359
65.0	260	70.8	380
66.7	275	72.4	398
67.2	289	73.3	421

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
7.82	5.07	12.43	76.8
8.32	11.36	13.38	94.2
8.80	18.12	13.70	101.5
9.34	25.25	15.21	129.7
9.98	34.52	16.80	159.3
10.61	44.60	18.69	193.7
11.52	60.9		

Jones and Shah, 1913

%	spontan. crystall.	%	spontan. crystall.
0	-0.6	37.46	-0.1
3.02	-1.9	37.74	0.9
7.56	-3.5	39.16	8.1
14.16	-6.8	40.34	12.1
17.35	-8.2	41.77	20.2
23.75	-11.3	42.85	26.2
30.55	-15.3	43.66	30.5
33.15	-16.4	44.40	34.9
34.55	-11.9	47.03	52.6
36.46	-3.5	47.97	60.4
37.09	-2.3	53.02	107.0
37.36	-1.3		

Properties of phases

Density

Bischof, 1850

%	d	%	d
	18.075°		
2	1.0123	18	1.1334
4	.0264	20	.1502
6	.0407	22	.1672
8	.0553	24	.1848
10	.0703	28	.2212
12	.0853	30	.2401
14	.1011	35	.2895
16	.1171		

Kremers, 1855 and 1858

%	d	%	d
	19.5°		
0	0.998	25	1.205
5	1.035	30	.254
10	.073	35	.307
15	.114	40	.364
20	.157	45	.428

t	d		
19.5	1.1340	1.2455	1.3385
0	1.1398	1.2538	-
40	.1252	.2350	1.3266
60	.1145	.2232	.3141
80	.1022	.2101	.3004
100	.0884	.1960	.2859

Fouqué, 1867

%	d 0°	d 22°
0.7	1.0049	1.0028
3.2	.0264	.0243
18.8	.2001	.1927

Kohlrausch, 1879

%	d 15°
5	1.0357
10	.0741
20	.1583
30	.2553
36	.3198

Schneider, 1881

N	d	N	d
	18°		
0	0.99862	0.9833	1.08002
0.0992	1.00700	1.973	.15982
.1976	.01527	4.032	.3226
.505	.04025		

Traube, 1885

c	%	d
	15°	
10	9.34	1.0703
20	17.57	1.1382

Röntgen and Schneider, 1886

%	d
18°	
7.68	1.1156
13.93	.1041
14.90	.1156

Wegner, 1889

%	d
20°	
0	0.998230
2.1644	1.014061
5.013	.035104
8.5659	.062645
19.8240	.158536

Jahn, 1891

c	d
20°	
0	0.9982
12.859	1.0874
20.671	1.1423

Sentis, 1897

mol%	t	d
1	17.9	1.0445
2	17.6	.0895
3	19.2	.1333
5	12.5	.2158
8	21.3	.3263

Oppenheimer, 1898

%	d
20°	
11.03	1.0815
20.49	.1629
28.83	.2407
37.97	.3449

Taylor and Ranken, 1903-04

N	d		
	0°	15°	25°
1	1.0858	1.0831	1.0804
2	.1692	.1662	.1623
3	.2521	.2453	.2413

Loewenfeld, 1905

%	d
15°	
0	0.9993
10.9	1.0835
20.5	.1659
28.7	.2458
36.3	.3256

Chéneveau, 1907

%	d	%	d
	15°		
0	0.9991	24.64	1.2064
4.76	1.0344	27.97	.2398
9.29	.0700	31.08	.2727
13.46	.1047	34.07	.3062
17.38	.1389	36.88	.3388
21.04	.1734		

Getman , 1907

%	d	%	d
	18°		
0	0.9986	17.61	1.1355
4.90	1.0199	21.77	.1484
9.44	.0585	25.83	.1614
13.67	.0970		

Getman and Wilson, 1908

%	d	%	d
	20°		
0	0.998	20	1.159
5	1.037	25	.207
10	.070	30	.256
15	.116		

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.00294	1.0	1.0817
.1	.00712	2.0	.1628
.2	.01557	4.0	.3217
.5	.0405		

Baxter, Boylston and al., 1911

%	d	%	d
	25°		
0	0.99707	4.2999	1.02821
0.3786	.99971	.4552	.02941
.9874	1.00396	7.1636	.05006
1.4297	.00718	.8856	.05543
.4978	.00776	8.8542	.06317
2.9364	.01829	9.7228	.06977
3.6692	.02373	10.7475	.07881
4.1387	.02704	21.017	.16718
.2164	.02770		

Clausen, 1912

M	d		
	6°	18°	30°
0.501	1.0424	1.04025	1.03673
1.000	.08402	.08110	.07705
2.005	.1666	.16247	.1575
4.022	.3275	.32327	.3206

Buchanan, 1912-13

N	d	N	d
	19.5		
0.002	0.998530	0.06	1.003642
.004	.998696	.12	.008883
.008	.999023	.25	.019241
.016	.999676	.5	.039583
.03	1.001006		

m	d
	19.5°
1	1.0790
2	.1526
3	.2200
4	.2832
5	.3403

m	t	d
	sat.sol.	
5.7250	23.4	1.3746

Grufki, 1913

molarity	d
	18°
0	0.99862
0.5018	1.04052
1.006	.08188
2.018	.16391
4.046	.32307

Herz, 1914 and 1917

N	d	N	d
	25°		
0.490	1.0372	2.938	1.2370
0.979	.0770	3.917	.3148
1.958	.1579	4.710	.3797

Gropp, 1915

t	d			
	0.498N	0.999N	2N	4.062N
		18°		
0	1.0421	1.0840	1.1680	1.3327
18	.0398	.0809	.1624	.3238
48	.0295	.0692	.1492	.3078
78	.0137	.0530	.1318	.2877
100	0.9995	.0388	.1175	.2732

Baxter and Wallace, 1916

%	d	%	d	%	d
0°		25°		50.04°	
2.69	1.00700	2.67	1.01627	2.66	1.01984
5.31	.02600	5.26	.03554	5.24	.03981
10.27	.06395	10.17	.07386	10.15	.07952
23.22	.17511	23.02	.18625	22.88	.19450
40.27	.35920	39.96	.37090		

Stocker, 1920

%	t	d
0	18.5	0.9985
6.97	19.3	1.0509
14.45	17.6	.1125
21.19	21.19	.1740
28.39	28.39	.2446

Manchot, Jahrstorfer and Zepter, 1924

c	d	c	d
25°			
13.211	1.0891	13.973	1.0866
25.590	.1752	25.947	.1750
49.864	.3380	51.893	.3459

de Block, 1925

%	d	%	d
11.5°			
0	0.9996	19.7	1.1620
5.6	1.0420	29.4	.2590
10.2	1.0836	36.2	.3328

Rakshit, 1925

%	d
20°	
1	1.00541
5	.03385
10	.06861
30	.20518
50	.33867

Hrynakowski, 1927

sat.sol. 50.50° d=1.4221

Scott and Frazier, 1927

40.57% d^{25}=1.37937

Flöttmann, 1928

%	t	d
38.59	15	1.3597
39.73	20	.3701
40.71	25	.3794

Scott and Durham, 1930

%	t	d
sat.sol.		
35.08	0.00	1.3237
44.58	35.00	.3941
44.95	50.21	.4160
50.37	91.95	.459

Hölemann and Kohner, 1931

m	d	m	d
25°			
0	0.99707	3.2144	1.23229
1.4872	1.11409	4.0556	1.28410
2.3613	1.17601	5.1138	1.34473
35°			
0	0.99406	3.2292	1.22748
1.4871	1.11012	5.1169	1.33909
2.1829	1.15922		
45°			
0	0.99024	3.2152	1.22192
1.4883	1.10544	5.1305	1.33368
2.2375	1.15795		

Thomas and Perman, 1932

%	d	%	d
	30°		
3.23	1.027	25.51	1.218
9.60	.077	31.16	.265
14.60	.123	32.10	.268
19.51	.166		

Jones and Talley, 1933

N	d	N	d
	25°		
0.001000	0.997163	0.499927	1.039157
.002000	.997234	.959172	.077301
.005000	.997491	.998357	.080570
.010005	.997912	1.999826	.162661
.020001	.998764	2.003090	.162797
.050001	1.001318	3.030933	.245986
.099899	.005542	3.749274	.303450
.199882	.013961		

Jones and Bickford, 1934

N	d	N	d
25°		0°	
0.00025	0.99709	0.00025	0.99989
.00036	.99710	.00036	.99990
.0005	.99711	.0005	.99991
.00075	.99713	.00075	.99994
.001	.99715	.001	.99996
.0016	.99721	.0016	1.00001
.002	.99725	.002	.00004
.005	.99746	.005	.00029
.01	.99793	.01	.00076
.02	.99877	.02	.00165
.05	1.00129	.05	.00430
.1	.00553	.1	.00869
.2	.01397	.2	.01746
.5	.03912	.5	.04356
1.0	.08061	1.0	.08635
2.0	.16255	2.0	.17014
3.0	.24328	3.0	.25218
3.75	.30331	3.75	.31286

Scott, Obenhaus and Wilson, 1934

%	d
35°	
10.452	1.07209
22.985	.18086
32.576	.27822
40.058	.36525

Gibson, 1935

%	d	%	d
	25°		
0.00	0.9979	26.47	1.2201
8.26	1.0586	32.88	.2884
13.79	.1034	39.02	.3594
20.00	.1579		

Hering, 1936

%	mol%	t	d
40.65	9.40	25	1.379
51.2	13.72	100	1.465

Jones and Stauffer, 1940

N	d	N	d
	0°		
0.00025	0.99989	0.1	1.00869
.0005	.99994	.2	.01747
.001	.99995	.5	.04343
.002	1.00004	1	.08629
.005	.00032	2	.17007
.01	.00077	2.98978	.25120
.02	.00165	3.750	.31269
.05	.00435		

Guillaume, 1946

%	d
20°	
12.70	1.0962
39.17	1.3657

de Lannoy, 1895

t	relative volume			
	4%	10%	20%	30%
0	1.00000	1.00000	1.00000	1.00000
10	.00072	.00160	.00260	.00272
20	.00267	.00391	.00578	.00577
30	.00595	.00710	.00952	.00913
40	.01019	.01085	.01353	.01274
50	.01438	.01532	.01805	.01669
60	.01903	.02040	.02315	.02103
70	.02433	.02621	.02888	.02562
80	.03042	.03281	.03500	.03085
90	-	-	-	.03632

Pohl, 1852

%	π (water=1)	%	π (water=1)
	13.9°		
0	1.000	26.27	0.830
9.60	0.906	33.41	0.709
17.61	0.838		

Röntgen and Schneider, 1886

%	π (water=1)
18°	
7.68	0.930
13.93	.872
14.90	.864

Freyer, 1931

%	π	t	π
20°		30%	
1	45.74	20	33.70
6	43.50	25	33.40
10	41.80	30	33.17
20	37.74	40	32.90
30	33.70	50	32.85
40	29.80		

Thomas and Perman, 1934

%	π	%	π
	30°		
3.23	43.0	25.51	33.7
9.60	40.1	31.16	31.7
14.60	37.8	32.10	31.4
19.51	35.8		

Scott, Obenhaus and Wilson, 1934

%	π	%	π
	35°		
10.452	38.54	32.576	31.72
22.985	34.68	40.058	29.51

Gibson, 1935

%	π (1-1000 bars)	%	π (1-1000 bars)
	25°		
0.00	39.35	26.47	31.72
8.26	36.94	32.88=	29.91
13.79	35.31	39.02	28.17
20.00	33.49		

Viscosity and surface tension

Schneider, 1881

N	η (water=1)	N	η (water=1)
	18°		
0.0992	0.9924	0.9833	0.9533
.1976	.9887	1.973	.9285
.505	.9738	4.032	.9599

Taylor and Ranken, 1903-04

N	η (water°=1) 0°	15°	25°
1	0.911	0.601	0.483
2	.837	.585	.477
3	.815	.582	.486

Getman , 1907

%	η	%	η
	18°		
0	1058	17.61	1089
4.90	1053	21.77	1106
9.44	1056	25.83	1120
13.67	1079		

c	η 10°	30°	50°
0	1310	800	550
15.99	1210	780	570
26.60	1180	780	590
46.64	1190	830	630

Lorenz, 1912

N	t	η
17.93	745	1480
16.73	775	1340
16.52	805	1190

Herz, 1914 and 1917

N	η	N	η
25°			
0	895	2.938	871
0.490	885	3.917	905
0.979	866	4.710	950
1.958	860		

Hrynakowski, 1927

sat.sol. 50.50° η=720

Tammann and Rabe, 1928

%	η 0°	10°	30°	75°
10.63	1578.4	1205.8	787.6	400.7
19.00	1488.1	1179.3	786.4	419.0
26.00	1455.7	1146.5	801.7	442.6
32.14	1435.7	1165.2	815.7	464.2
37.25	-	1202.4	846.8	477.3

Jones and Talley, 1933

N	η (water=1)	N	η (water=1)
25°			
0.001000	1.00010	0.499927	0.98180
.002000	.00010	.959172	.96885
.005000	.00009	.998357	.96798
.010005	0.99998	1.999826	.95756
.020001	.99972	2.003090	.96953
.050001	.99874	3.030933	.99177
.099899	.99678	.749274	.997074
.199882	.99287		

Jones (Gr.) and Stauffer, 1940

N	η (water=1)	N	η (water=1)
0°			
0.00025	1.00007	0.1	0.98635
.0005	.00005	.2	.97284
.001	0.99992	.5	.93649
.002	.99985	1.0	.88866
.005	.99950	2.0	.82882
.01	.99896	2.98978	.80363
.02	.99760	3.750	.80234
.05	.99330		

Traube, 1885

c	%	σ
15°		
10	9.34	73.62
20	17.57	74.92

Sentis, 1897

mol%	t	σ	mol%	t	σ
0	25.1	72.3	3	19.2	75.7
0	13.5	74.0	5	12.5	77.7
1	17.9	74.2	8	21.3	78.8
2	17.6	75.0			

Loewenfeld, 1905

%	σ
15°	
0	76.16
10.9	73.16
20.5	71.14
28.7	75.56
36.3	77.01

Stocker, 1920

%	σ
18°	
0	72.65
6.97	73.62
14.45	74.37
21.19	75.63
28.39	77.00

de Block, 1925

%	σ	%	σ
11.5°			
0	73.78	19.7	76.21
5.60	74.64	29.4	77.60
10.20	75.02	36.2	78.70

Hrynakowski, 1927

sat.sol. 50.50° σ= 7346

Wegner, 1889

%	n Li	n Hα	n D	n Hβ	n Hγ
			20°		
0	1.33082	1.33116	1.33298	1.33715	1.34038
2.1644	.325747	.333746	.335628	.339716	.34299
5.013	-	.337134	.339066	-	-
8.5659	-	.341573	.343570	-	-
19.8240	-	.356673	.358884	-	-

Bender, 1890

N(15°)	n Hα	n D	n Hβ	n Hγ
		20.5°		
0	1.33100	1.33297	1.33703	1.34021
1	.34428	.34628	.35093	.35458
2	.35688	.35907	.36413	.36823
3	.36917	.37156	.37697	.38138

Jahn, 1891

c	n Hα	n D	n Hβ
		20°	
0	1.3315	1.3332	1.3375
12.859	.3457	.3477	.3524
20.671	.3543	.3565	.3615

Borgesius, 1895

%	t	n_D (water=1)	%	t	n_D (water=1)
1.80	18.4	1.000216	12.82	21.9	1.001746
3.57	18.1	.000442	12.82	18.3	.001760
6.86	18.4	.000880	22.60	18.3	.003515

Kremers, 1857

%	n_D
17°	
0	1.3332
23.20	.3626
37.70	.3859

Wagner, 1903

c	n_D	c	n_D
	17.5°		
0	1.33320	10.183	1.34500
0.323	.33358	.516	.34537
.646	.33397	.849	.34575
.969	.33435	11.182	.34612
1.293	.33474	.515	.34650
.617	.33513	.848	.34687
.941	.33551	12.181	.34724
2.266	.33590	.514	.34761
.591	.33628	.847	.34798
.916	.33667	13.180	.34836
3.242	.33705	.513	.34873
.568	.33743	.846	.34910
.895	.33781	14.179	.34947
4.222	.33820	.512	.34984
.550	.33858	.845	.35021
.878	.33896	15.178	.35058
5.208	.33934	.511	.35095
.539	.33972	.844	.35132
.870	.34010	16.177	.35169
6.210	.34048	.510	.35205
.532	.34086	.844	.35242
.863	.34124	17.178	.35279
7.194	.34162	.512	.35316
.526	.34199	.846	.35352
.858	.34237	18.180	.35388
8.190	.34275	.514	.35425
.522	.34313	.848	.35461
.854	.34350	19.182	.35497
9.186	.34388	.516	.35533
.518	.34426	.850	.35569
.850	.34463	20.184	.35606

Jones , 1904 and Jones and Getman, 1904

N	n_D	N	n_D
	0°		
0.099	1.32648	1.598	1.34710
.199	.32798	.998	.35302
.399	.33069	2.397	.35824
.799	.33634	3.196	.36794
1.199	.34174		

Chéneveau, 1907

%	n_D	%	n_D
	15°		
0	1.3334	24.64	1.3667
4.76	.3392	27.97	.3718
9.29	.3450	31.08	.3769
13.46	.3505	34.07	.3821
17.38	.3560	36.88	.3871
21.04	.3615		

Getman and Wilson, 1908

%	n_D	%	n_D
	20°		
0	1.33298	20	1.35875
5	.33900	25	.36628
10	.34515	30	.37418
15	.35178		

Heydweiller, 1909

N	n_D
18°	
0	1.33327
0.5	.34011
1.0	.34658
2.0	.35967
4.0	.38409

Baxter, Boylston and al., 1911

%	n_D	%	n_D
	25°		
0	1.33246	4.2999	1.33763
0.3786	.33292	4.4552	.33777
0.9874	.33370	7.1636	.34115
1.4297	.33408	7.8856	.34200
1.4978	.33425	8.8542	.34327
2.9364	.33596	9.7228	.34435
3.6692	.33684	10.7475	.34575
4.1387	.33743	21.017	.35984
4.2164	.33749		

John, 1891

c	(α) magn.
20°	
0	1
12.859	1.4153
20.671	1.4160

Grufki, 1913

N	n H_α	H_β	H_γ
		18°	
0	1.33140	1.33732	1.34050
0.5018	.33818	.34444	.34782
1.006	.34479	.35136	.35492
2.018	.35765	.36485	.36883
4.046	.38205	.39046	.39525

Flöttmann, 1928

%	t	n_D
38.59	15	1.38989
39.73	20	.39115
40.71	25	.39221

Hölemann and Kohner, 1931

N	n_{He}	N	n_{He}	N	n_{He}
25°		35°		45°	
0	1.33266	0	1.33142	0	1.32995
1.4872	.35158	1.4871	.35017	1.4883	.34857
2.3613	.36133	2.1829	.35789	2.2375	.35684
3.2144	.37008	3.2292	.36857	3.2152	.36682
4.9556	.37803	5.1169	.38554	5.1305	.38387
5.1138	.38721				

Guillaume, 1946

%	n 5780 Å	*(α) magn. $\cdot 10^6$ 5780
	20°	
0	-	3.974
12.70	1.3493	4.214
39.17	1.3905	4.646

in radians, gauss, cm

Oppenheimer, 1898

%	(α) magn.
20°	
11.03	1.41
20.49	1.41
28.83	1.40
37.97	1.42

Okazaki, 1933

%	Verdet's constant (3441 Å)
28°	
7.37	0.04927
12.75	.05317
17.19	.05656
21.41	.06042
27.18	.06622
34.47	.07305

Okazaki, 1936

%	Verdet's constant (D)
25°	
0	0.013075
2.102	.01339
2.147	.01341
4.154	.01370
4.251	.01375
4.96	.01383
7.19	.01421
7.82	.01431
9.04	.01452
9.43	.01459
11.20	.01489
14.71	.01549
17.27	.01600
20.73	.01661
22.82	.01710
26.16	.01773
31.51	.01889
37.40	.02029

Okazaki, 1951

%	Verdet's constant (D)
35°	
0	0.013047
14.66	.01543
17.78	.01594
20.97	.01660
25.59	.01757
31.96	.01900
36.90	.02014

Kohlrausch, 1879

%	κ	$\tau \cdot 10^4$
	18°	
5	464	207
10	925	195
20	1899	178
30	2912	165
36	3495	155

Dennhardt, 1899

N	λ 0°	λ 10°	λ 18°
0.5	72.9	91.7	106.0
1	70.1	88.6	103.0
2	66.3	84.1	97.5
3	63.9	80.1	93.0
4	61.6	75.7	88.0

Taylor and Ranken, 1903-04

N	λ
	0°
1	68.3
2	67.6
3	65.8

Jones , 1904 and Jones and Getman, 1904

N	λ	N	λ
		0°	
0.099	69.50	1.598	64.60
.199	69.00	1.998	64.20
.399	67.14	2.397	60.10
.799	66.95	3.196	57.64
1.199	65.00		

Heydweiller, 1909

N	λ	N	λ
		18°	
0.05	118.9	1.0	100.6
.1	112.8	2.0	95.5
.2	109.7	4.0	85.1
.5	104.5		

Sloan, 1910

N	λ	N	λ
	0°		
0.08	74.65	1.34	68.63
.17	72.50	2.69	67.30
.34	70.10	3.53	65.70
.67	69.35		

Clausen, 1912

N	λ		
	6°	18°	30°
0.501	79.34	103.47	128.4
1.000	77.58	99.87	123.1
2.005	75.72	95.46	115.5
4.022	70.02	85.84	101.77

N	$\tau \cdot 10^5$		
	6°	18°	30°
0.501	3978	5185	6452
1.000	7758	9987	12310
2.005	15180	19140	23190
4.022	28160	34520	40930

Lorenz, 1912

N	t	λ
17.93	745	76.96
16.73	775	95.60
16.52	805	113.19

Gropp, 1915

t	κ			
	0.498N	0.999N	2N	4.062N
0	337.8	667.1	1321	2515
18	513.0	990.2	1911	3467
48	836.5	1582	2947	5103
78	1168	2202	4023	6724
100	1437	2659	4797	7735

Scatchard and Prentiss, 1933

m	κ	m	κ
	10°		
0.001469	1.5	0.19939	190.7
.004512	4.7	.28251	264.6
.005370	5.7	.46754	422.2
.014042	14.7	.69490	606.5
.020343	21.1	.72135	627.4
.046983	47.8	1.07251	899.0
.096075	95.2	1.29601	1060.9

Jones and Bickford, 1934

N	λ	
	25°	0°
0.00025	150.16	82.57
.00036	149.87	82.43
.0005	149.55	82.26
.00075	149.12	82.04
.001	148.72	81.87
.0016	148.02	81.48
.002	147.64	81.32
.005	145.47	80.20
.01	143.15	79.06
.02	140.26	77.61
.05	135.44	75.25
.1	131.19	73.23
.2	126.59	71.22
.5	120.35	69.08
1.0	115.46	68.15
2.0	109.37	67.39
3.0	103.55	65.85
3.75	98.70	64.01

Thermal constants

Faasch, 1911

N	U
18°	
0.486	0.931
0.974	.880
1.950	.779
3.925	.638

Randall and Rossini, 1929

m	U	m	U
25°			
0.00	0.9979	0.20	0.9698
.01	.9964	.35	.9501
.02	.9950	.50	.9314
.05	.9906	.75	.9021
.10	.9835	1.00	.8750

Bender and Kaiser, Jr., 1954

molality	U 25°	U 30°
0.9985	0.8746	0.8764
1.9960	.7838	.7850
2.9945	.7133	.7142
3.9795	.6561	.6576
4.9985	.6082	.6095
5.5245	.5875	.5885

Scholz, 1892

N	Q diss (cal/g)
0.0625	-48.83
.125	48.82
.25	48.72
.5	47.62
1.0	46.04
2.0	43.08
4.0	38.99

Wüst and Lange, 1923 and 1924

N	Q diss integral	N	Q diss integral
0.662	-4821	0.330	-4821
1.328	4740	0.996	4661
1.975	4654	1.65	4482
2.617	4579	2.30	4343
3.244	4498	2.92	4187
3.868	4430	3.54	4077
4.475	4363	4.18	3961
5.087	4300	4.79	3885
5.686	4244	5.39	3770
6.295	4188	5.99	3704
6.874	4136	6.57	3611
7.450	4087	7.16	3543
8.039	4036	7.75	3465
8.600	3990	8.31	3393

Kapustinskii and Ruzavin, 1955

%	c	%	c
25°			
Dt=4°			
1.97	1445	21.3	1338
4.83	1429	24.9	1321
9.38	1403	28.2	1298
13.6	1383	31.5	1277
17.6	1359		
Dt=2°			
4.97	1427	20.6	1341
9.73	1400	31.2	1269
14.7	1371		

$c.10^6$ = heat conductivity coefficient

Water (H_2O) + Potassium iodide (KI)

Heterogeneous equilibria

Tammann, 1885

t	p 0%	11.57%	26.45%	39.79%	49.07%
32.70	37.1	36.3	34.6	32.7	30.6
41.59	60.2	59.0	56.2	52.8	48.9
46.46	77.5	76.0	72.4	67.4	62.4
51.16	98.0	95.7	91.2	84.6	77.9
53.01	107.3	104.8	99.8	93.2	85.9
57.40	132.4	-	123.6	115.5	106.5
61.32	158.8	155.3	147.4	137.1	125.8
71.64	250.9	245.3	232.7	215.7	198.2
75.51	295.5	289.3	274.3	253.9	230.3
79.15	343.4	335.0	316.9	293.5	269.6
83.62	410.9	401.9	380.7	351.9	323.2
90.18	529.6	518.2	490.7	451.2	415.4
93.59	601.8	589.2	558.4	511.7	471.0
96.76	676.2	-	628.5	578.5	530.4
100.66	778	759.0	720.0	663.8	-

%	p	%	p
	100°		
9.26	744.4	52.63	558.7
18.79	725.3	53.61	552.1
23.50	712.4	57.44	528.4
31.81	682.8	57.94	525.2
35.38	669.0	62.84	476.6
38.12	656.1	64.68	459.8
41.71	638.3	66.70	438.7
49.21	592.6		

Dieterici, 1891

%	p	%	p
	0°		
0	4.620	49.90	3.704
24.92	4.316	55.46	3.474
39.90	4.012		

Tower, 1908

%	p	%	p
	0°		
0	4.579	8.61	4.513
1.32	.564	19.70	.433
2.67	.555	34.53	.178
4.42	.541	65.89	3.472

Boswell and Cantelo, 1922

N	Dp/N	N	Dp/N
	23°		
1.000	0.029	5.000	0.235
2.000	.068	6.500	.318
3.500	.0147		

Furutani, 1927

t	p 1.60%	3.10%	5.90%	9.80%	11.49%
100	0.926	0.90	0.83	0.765	0.722
105	1.17	1.10	1.01	.921	.861
110	.37	.31	.18	.12	1.07
115	.60	.53	.38	.30	.25
120	.87	.78	.67	.50	.47
125	2.22	2.10	.92	.74	.70
130	.55	.44	2.23	2.05	.94
135	.95	.83	.55	.40	2.29
140	3.45	3.28	.96	.78	.65
145	.92	.78	3.42	3.15	3.04
150	4.49	4.32	.94	.62	.45
155	5.14	.92	4.49	4.15	.93
160	.85	5.58	5.12	.71	4.50
165	6.62	6.38	.76	5.38	5.14
170	7.40	7.20	6.46	6.05	.81
175	8.34	8.02	7.30	.80	6.54
180	9.45	8.98	8.19	7.58	7.35
190	10.50	10.05	9.15	8.45	8.19
195	11.65	11.09	10.18	9.36	9.09
200	12.85	12.30	11.20	10.38	10.00

t	p 14.01%	18.42%	22.62%	25.13%
100	0.712	0.605	-	-
105	.859	.720	-	-
110	.991	.902	-	-
115	1.12	1.01	-	-
120	.36	.22	1.12	1.12
125	.57	.41	.30	.31
130	.83	.62	.51	.51
135	2.16	.90	.76	.77
140	.48	2.20	2.04	2.04
145	.82	.52	.32	.29
150	3.26	.95	.70	.61
155	.72	3.40	3.15	.97
160	4.26	.90	.60	3.29
165	.85	4.44	4.10	.76
170	5.47	5.00	.62	4.24
175	6.11	5.60	5.17	.75
180	.87	6.25	.76	5.31
185	7.74	6.97	6.45	.92
190	8.60	7.75	7.20	6.60
195	9.57	8.55	.85	7.30
200	10.41	9.40	8.72	8.10

%	p	%	p
	sat.sol.		
18.42	0.606	23.09	2.97
18.90	.723	23.58	3.30
19.49	.861	23.98	3.76
19.80	.998	24.57	4.24
20.20	1.12	25.13	4.68
20.62	.30	25.58	5.11
21.01	.52	26.32	5.71
21.32	.77	27.17	6.32
21.79	2.04	27.47	6.92
21.98	.30	28.33	7.58
22.42	.62		

Pearce, Taylor and Bartlett, 1928; Pearce and Nelson, 1932

m	p	m	p
	25°		
0.0	23.752	3.5	21.100
.1	23.686	4.0	20.681
.2	23.620	4.5	20.253
.4	23.486	5.0	19.822
.6	23.384	5.5	19.392
.8	23.211	6.0	18.966
1.0	23.072	6.5	18.529
1.5	22.702	7.0	18.082
2.0	22.316	7.5	17.635
2.5	21.922	8.0	17.190
3.0	21.516	8.5	16.750
8.938 std	16.370		

Lannung, 1934

m	p	m	p
	18°		
0.2636	15.35	1.047	14.97
.2724	15.34	.130	14.92
.2887	15.33	.262	14.85
.3320	15.32	.466	14.74
.3508	"	.684	14.63
.4257	15.28	.945	14.50
.4817	15.24	3.305	13.78
.5585	15.20	5.627	12.50
.6665	15.15	sat.sol.	12.90
.767	15.09		

Schlamp, 1858

%	b.t.	%	b.t.
0	100	10.09	100.656
1.42	100.101	13.49	100.911
4.13	100.256	15.39	101.076
7.28	100.406		

Gerlach, 1886

%	b.t.	%	b.t.
0	100	57.27	110
13.04	101	59.18	111
23.08	102	60.78	112
31.03	103	62.26	113
37.50	104	63.64	114
42.53	105	64.91	115
46.52	106	66.10	116
49.88	107	67.21	117
52.71	108	68.25	118
55.16	109	68.75	118.5

Buchanan, 1899

m	b.t.	m	b.t.
0.43	100.01	1.43	101.02
.55	100.11	.54	101.12
.66	100.21	2.00	101.62
.77	100.31	.42	102.15
.87	100.41	3.19	103.14
.97	100.51	.90	104.15
1.08	100.61	4.90	105.67
.26	100.81		

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
2.72	100.16	22.63	101.86
4.00	100.24	25.19	102.19
5.81	100.34	25.93	102.19 sic
6.57	100.41	28.09	102.57
9.75	100.64	30.10	102.80
11.45	100.74	32.25	103.24
12.98	100.89	39.85	104.77
15.99	101.12	43.55	105.67
16.52	101.22	46.16	106.50
19.70	101.53	48.12	107.17
20.49	101.61	51.21	108.19

Johnston, 1907

%	b.t.	%	b.t.
4.66	100.351	30.44	102.677
6.55	100.394	33.4	103.184
13.15	100.866	36.87	103.758
14.44	101.010	39.67	104.360
19.28	101.350	42.11	104.901
23.08	101.721	45.9	105.955
26.53	102.146	49.5	106.709

Jablezynski and Kon, 1923

m	b.t.	m	b.t.
0.4191	100.390	1.0056	100.983
.5782	100.549	.1432	101.130
.7114	100.682	.2778	101.273
.8709	100.846		

Robinson and Stokes, 1949

m	osmotic coefficient
25°	
0.1	0.932
.2	.922
.3	.918
.4	.917
.5	.917
.6	.918
.7	.919
.8	.922
.9	.924
1.0	.926
.2	.931
.4	.937
.6	.943
.8	.950
2.0	.957
2.5	.974
3.0	.990
3.5	1.006
4.0	.021
	.032

Robinsons, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1042	0.1035	2.157	2.066
.1053	.1047	.555	.445
.1708	.1691	.600	.494
.3109	.3068	.807	.689
.5536	.5426	3.111	.969
.7025	.6852	.131	.992
1.022	.9892	.470	3.324
.217	1.175	.498	.355
.224	.179	.666	.499
.278	.232	.931	.748
.715	.647	4.400	4.198
.857	.779	.81	.581

m_1 = molality of potassium chloride

m_2 = " " " iodide

Kremers, 1856

%	f.t.	%	f.t.
55.87	0	65.36	80
58.24	20	66.22	100
61.35	40	sat.sol.	119
63.69	60		

Rüdorff, 1862

%	f.t.
3.62	-0.8
9.37	-2.2
20.78	-5.55
36.77	-12.25

Gérardin, 1865

%	f.t.	%	f.t.
57.36	8	62.26	46
58.09	13	63.24	55
58.73	18	64.03	62
59.66	27		

Mulder, 1866

%	f.t.	%	f.t.
56.10	0	62.69	49
56.79	5	64.30	65
60.49	31	65.60	78.5
60.99	35	69.01	118.4

Guthrie, 1876

%	f.t.	%	f.t.
10	-2.2	52.07	-22 E
20	-5.1	55.93	0
30	-9.0	58.9	+20
40	+4.4	61.4	+40

De Coppet, 1883

%	f.t.	%	f.t.
51.73	-22.65	61.58	42.3
51.59	-22.35	62.06	45.75
52.62	-16	62.62	51.8
53.76	-11.35	62.83	55.05
54.63	-5.9	63.52	60.55
55.68	0.0	64.07	65.0
56.54	3.25	64.73	71.1
57.22	9.55	65.76	79.75
57.83	12.75	66.05	81.6
57.95	12.9	66.70	86.35
58.89	21.05	67.28	93.5
59.44	25.6	68.36	100.7
59.94	29.1	68.63	110.2
61.04	37.3	64.98	113.7

Tilden and Shenstone, 1884

%	f.t.	%	f.t.
70.15	124	72.57	144
71.37	133	75.63	175

Etard, 1894

%	f.t.	%	f.t.
51	-21	66.9	96
52.2	-19	70.6	150
53.2	-15	70.9	151
54.5	-9	71.6	175
56.9	0	72.7	176
59.3	21	73.8	190
60.8	44	74.5	193
64.3	72	75.7	213
64.8	78		

Jones, 1904 and Jones and Getman, 1904

N	f.t.	N	f.t.
0.068	-0.223	1.641	- 6.120
0.137	-0.473	2.188	- 8.500
0.273	-0.945	2.735	-10.900
0.547	-1.873	3.282	-13.600
1.094	-3.880	3.829	-16.550

Meusser, 1905

%	f.t.	%	f.t.
22.5	-5	42.7	-14
25.6	-6	53.5	-10
29.9	-7	54	-4
34.0	-9.5	54.5	-5
39.3	-11.5	55.0	-1

Kremann and Kerschbaum, 1907

%	f.t.	%	f.t.
36.3	-12.0	49.0	-20.2 sic
39.7	-13.7	51.0	-21.5
41.6	-16.0	51.6	-22.4
43.1	-15.9	52.5	-21.3
43.5	-17.0	53.5	-14.4
45.1	-17.6	53.6	-15.3
46.3	-18.3	53.8	-14.2
46.4	-18.8	54.9	-8.9
47.4	-20.2	55.6	-4.0
48.7	-20.8		

Ceitlin, 1925

N	f.t.
5.86	10.2
6.11	20.0

Kordes, 1926

mol%	f.t.
0	0
10.5	-23 E
100	681

Scott and Frazier, 1927

f.t. = 25° 59.75 %

Kracek, 1931

mol%	f.t.	mol%	f.t.
17.08	78.1	22.91	181.7
18.46	102.4	24.08	199.3
19.51	121.6	24.28	201.5
20.265	134.9	24.53	206.0
21.35	153.7	25.01	213.3
21.33	154.3	28.56	230.4

Hering, 1936

wt%	mol%	f.t.	wt%	mol%	f.t.
51.35	10.28	-21.7	63.75	16.03	60
56.05	12.16	0	65.6	17.15	80
59.15	13.59	20	67.35	18.29	100
61.7	14.87	40			

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
68.0	108	80.2	327
68.7	115	81.7	330
69.2	131	82.1	335
71.2	147	83.2	351
72.3	163	82.6	371
73.0	175	83.1	375
74.3	198	84.6	378
76.2	219	84.9	383
78.0	256	86.3	402
79.4	279	86.7	417
79.9	203	86.7	418
80.2	301	88.9	450

Eddy and Menzies, 1940

mol%	f.t.
12.58	6.2
12.89	11.2

Keevil, 1942

mol%	f.t.	P	mol%	f.t.	P
		sat.sol			
30.1	306.3	39.91	57.9	513.0	87.46
39.3	388.0	73.85	62.7	538.0	82.03
52.0	476.5	92.09	67.4	563.0	74.71
56.8	506.5	90.01			

Jones and Shah, 1913

%	spontan. crystalliz.	%	spontan. crystalliz.
0	-0.6	59.32	19.5
1.90	-1.3	59.53	21.1
9.24	-2.8	61.19	34.2
13.27	-3.9	61.43	36.2
17.70	-4.8	62.16	43.0
21.38	-6.3	62.85	48.9
27.46	-7.8	63.48	55.0
55.92	-2.9	64.24	63.1
57.58	+7.0	65.11	71.7
57.97	9.4	66.13	82.7
58.62	14.6		

Properties of phases.

Density

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0132	20	1.1662
4	.0284	24	.2063
6	.0440	28	.2492
8	.0600	32	.2716
10	.0764	35	.3313
12	.0933	40	.3962
14	.1107	44	.4667
16	.1286	50	.5434
18	.1472	55	.6266

Grassi, 1851

%	t	d
0	13.4	0.9994
58.67	15.5	1.694

Kremers, 1855 and 1859

%	d	%	d
	19.5°		
5	1.036	35	1.329
10	.076	40	.394
15	.118	45	.467
20	.164	50	.543
25	.216	55	.633
30	.269	60	.731

t	d			
19.5°	1.1836	1.3423	1.5119	1.6794
0	1.1905	1.3527	1.5252	-
40	.1737	.3296	.4967	1.6624
60	.1620	.3158	.4810	.6451
80	.1490	.3007	.4644	.6273
100	.1939	.2845	.4470	.6094

Schiff, 1858

%	d	%	d	%	d
		15°			
1	1.008	25	1.229	48	1.510
2	.016	26	.240	49	.526
3	.024	27	.251	50	.538
4	.032	28	.263	51	.551
5	.040	29	.277	52	.564
6	.048	30	.287	53	.577
7	.057	31	.299	54	.589
8	.064	32	.310	55	.603
9	.073	33	.323	56	.617
10	.082	34	.335	57	.629
11	.091	35	.348	58	.640
12	.109	36	.360	59	.654
13	.110	37	.373	60	.666
14	.118	38	.386	61	.681
15	.127	39	.399	62	.694
16	.136	40	.410	63	.704
17	.145	41	.424	64	.716
18	.154	42	.437	65	.727
19	.165	43	.449	66	.738
20	.176	44	.461	67	.749
21	.187	45	.474	68	.766
22	.197	46	.487	69	.778
23	.208	47	.498	70	.788
24	.219				

Schiff, 1858 and 1859

%	d	%	d	%	d
		21°			
0	0.9980	21	1.1785	41	1.4197
1	1.0056	22	.1887	42	.4344
2	.0132	23	.1992	43	.4493
3	.0208	24	.2098	44	.4643
4	.0285	25	.2205	45	.4797
5	.0364	26	.2312	46	.4954
6	.0444	27	.2420	47	.5113
7	.0525	28	.2531	48	.5276
8	.0607	29	.2644	49	.5442
9	.0690	30	.2759	50	.5610
10	.0773	31	.2874	51	.5780
11	.0856	32	.2992	52	.5954
12	.0941	33	.3113	53	.6131
13	.1027	34	.3237	54	.6496
14	.1115	35	.3364	55	.6685
15	.1205	36	.3494	56	.6879
16	.1297	37	.3627	57	.7076
17	.1390	38	.3765	58	.7278
18	.1486	39	.3907	59	.7484
19	.1583	40	.4052	60	.6312
20	.1683				

Fouqué, 1867

%	d	
	0°	22°
0.68	1.0045	1.0030
3.92	1.0284	1.0259
18.73	1.1587	1.1535

Kohlrausch, 1879

%	d
18°	
5	1.0363
10	.0762
20	.1679
30	.2730
40	.3966
50	.545
55	.630

Damien, 1881

%	d
20°	
0	0.99827
18.70	1.1562
23.44	.2180
39.70	.3920
44.22	.4475
55.12	.6290

Traube, 1885

c	d
15°	
10	1.0737
20	1.1440

Röntgen and Schneider, 1886

%	d
18°	
10.27	1.0794
19.70	1.1646

Wegner, 1889

%	d
20°	
2.9845	1.020295
5.0089	.035839
10.2725	.078430
15.2112	.120837
20.0390	.166542

Jahn, 1891

c	d
20°	
0	0.9982
17.535	1.1244
33.407	1.2378

Humburg, 1893

%	d
16°	
0	0.9990
15.4670	1.1258
39.0736	1.3835

Schönrock, 1893

%	d
16°	
0	0.999
32.4875	1.30238

Delaite, 1894

%	d	%	d
15°			
0.06	0.9992	2.04	1.0141
0.13	0.9997	4.02	1.0295
0.26	1.0006	7.80	1.0598
0.52	1.0027	14.78	1.1193
1.03	1.0066	25.15	1.2357

Sentis, 1897

mol %	t	d
1	18.8	1.0630
1	13.2	.0653
2	18.0	.1270
3	17.5	.1866
5	20.2	.2599
6.64	20.0	.3892
10	19.7	.5536

Gilbault, 1897

%	d	%	d
20°			
0	0.9982	25	1.216
5	1.036	30	1.269
10	1.076	40	1.394
15	1.118	50	1.543
20	1.164	60.15	1.734

Taylor and Ranken, 1903-04

N	d	N	d	N	d
0°		15°		25°	
1	1.1212	1	1.1188	1	1.1159
2	.2415	2	.2365	2	.2323
3	.3621	3	.3552	3	.3499

Forch, 1905

d^{18} (2.005 N) = 1.2365

Grüneisen, 1905

M	d	M	d
18°			
0.05	1.00470	1	.1188
.10	.01076	5.245	.6139
.20	.02285	5.6	.6547
.50	.05891		

Agerer, 1905

%	d
18.5°	
10.491	1.0749
22.784	.1938
48.578	.5228

Walden, 1906

c	%	t	d	τ.10⁴
94.05	56.32	0	1.670	-
102.70	59.54	25	.722	37

Chéneveau, 1907

%	d	%	d
18.5°			
0	0.9985	25.48	1.2209
9.65	1.0751	32.35	.2962
18.08	.1513	37.91	.3688
21.89	.1861		

Getman , 1907

%	d	%	d
18°			
0	0.9986	34.13	1.3189
9.33	1.0711	36.94	.3536
17.50	.1424	42.13	.4241
21.23	.1775	46.81	.4955
24.74	.2125	51.08	.5662
28.05	.2483	55.01	.6371
31.16	.2840	58.61	.7061

Getman and Wilson, 1908

%	d	%	d
20°			
0	0.998	25	1.218
5	1.038	30	.271
10	.078	35	.331
15	.120	40	.396
20	.166	50	.546

Tower, 1908

%	d	%	d
0°			
0	0.9999	19.70	1.1677
4.42	1.0410	34.53	.3315
8.61	.0677	65.89	.6605

Heydweiller, 1909

N	d	N	d
18°			
0.05	1.00471	1.0	1.11857
.1	.01078	2.0	.23698
.2	.02285	4.0	.47000
.5	.05872		

Baxter, Boylston and al., 1911

%	d	%	d
25°			
0	0.99707	3.1122	1.01986
0.4379	1.00045	5.1029	.03515
0.9151	.00382	8.5864	.06282
1.2259	.00601	25.604	.2200
2.3607	.01219		

Clausen, 1912

mol%	d 6°	d 18°	d 30°
0.4987	1.06105	1.05861	1.05488
0.9965	.12156	.11812	.11355
1.990	.24120	.23608	.23010
3.980	.47583	.46806	.45993

Buchanan, 1912-13

m	d	m	d
19.5°			
1	1.1128	5	1.4766
2	.2157	6	.5458
3	.3097	7	.6115
4	.3959	8	.6722

N	d 23.0°	d 19.5°
1	-	1.112528
0.75	-	.085308
.50	1.056113	.057205
.25	.027260	.028229
.12	.012568	.013951
.06	.005140	.005948
.03	.001366	.002156
.015	0.999528	.000268
.008	.998562	0.999220
.004	.998100	.998851
.001	-	.998494

m	t	d
sat.sol.		
8.9844	24.3	1.7222

Grufki, 1913

N	d
18°	
0	0.99862
0.5015	1.05897
1.015	.012038
2.009	.23863
4.015	.47169

Lübben, 1913

N	%	d
	18°	
0	0	0.99862
0.4922	7.81	1.05791
0.9752	14.47	.11567
1.934	25.95	.22978
3.878	43.86	.45597

Herz, 1914

N	d	N	d
	25°		
0	0.9971	3.888	1.4502
1.296	1.1504	5.184	.5991
1.944	.2261	6.200	.7213
2.592	.3013		

Baxter and Wallace, 1916

%	d	%	d	%	d
0°		25°		50.04°	
4.42	1.03406	4.42	1.03011	4.46	1.02062
8.57	.06829	8.60	.06328	8.69	.05339
16.28	.13647	16.35	.12962	16.49	.11893
34.54	.33532	34.77	.32516	35.06	.31288
55.69	.66049	56.06	.64833	56.48	.63486

Herz, 1917

N	d	N	d
	25°		
1.296	1.1504	3.888	1.4497
1.944	.2257	5.184	.5990
2.592	.3008	6.200	.7204

de Block, 1925

%	d
15°	
0	0.9991
14.8	1.1192
24.5	.2140
46.7	.4959
56.0	.6538

Rakshit, 1925

%	d	%	d
	20°		
1	1.00548	50	1.35273
5	.03427	60	.4231
10	.07020	70	.4930
30	.21156	80	.5598

Hrynakowski, 1927

sat. sol. 50.50° d= 1.7671

Scott and Frazier, 1927

25° 59.75% d= 1.72111

Herz and Hiebenthal, 1929

N	d	
	25°	70°
0.5	1.0573	1.0320
1.0	.1166	.0944
2.0	.2336	.2080
4.0	.4649	.4318

Scott and Durham, 1930

%	t	d
56.18	0.00	1.6673
61.18	35.00	.7380
62.84	50.13	.7600
66.69	92.18	.8056

Shibata and Hölemann, 1935

c	d	c	d	c	d
25°		35°		45°	
0	0.99707	0	0.99406	0	0.99024
1.0578	1.11696	1.0766	1.11483	1.0069	1.10265
1.7670	.19039	1.9513	.20364	1.9969	.20246
2.3858	.25045	3.4145	.33664	3.5646	.34266
3.0811	.31391	4.8063	.44789	4.8280	.44223
4.7566	.45115				

Gibson, 1934

c	d	c	d
	25°		
0	0.99708	49	1.34576
1	1.00484	64	.45087
4	.0261	64.415	.45377
8.877	.06109	80.785	.56778
9	.06197	81	.56927
16	.11208	81.002	.56928
16.040	.11236	83.759	.58840
25	.17619	90.25	.63335
25.052	.17656	93.940	.65885
36	.25414	95.00	.66617
36.010	.25420	95.028	.66637
48.414	.34164		

Scott, Obenhaus and Wilson, 1934

%	d
35°	
4.4120	1.02648
8.9594	.06226
16.415	.12576
34.901	.32072
58.262	.68149

Thomas and Perman, 1934

%	d	%	d
	30°		
6.32	1.048	31.02	1.279
12.41	.101	38.76	.379
19.86	.165	47.53	.476

Hering, 1936

%	mol %	t	d
51.35	10.28	-21.7	-
56.05	12.16	0	1.667
59.15	13.59	20	1.716
61.7	14.87	40	1.747
63.75	16.03	60	1.773
65.6	17.15	80	1.795
67.35	18.29	100	1.819

Forch, 1905

t	relative volume 0.501 N	1.002 N	2.005 N
		18°	
0	1.000000	1.000000	1.000000
5	.000444	.000861	.001494
10	.001177	.001974	.003112
15	.002155	.003278	.004858
20	.003391	.004773	.006753
25	.004829	.006444	.008764
30	.006448	.008265	.010877
35	.008287	.010242	.013117
40	.010290	.012380	.015476

Röntgen and Schneider, 1886

%	π (water = 1)
18°	
10.27	0.933
19.70	0.871

Gilbault, 1897

P	v/v_1	P	v/v_1
	30%	20°	
1	1.00000	150	1.99403
25	.99897	175	.99308
50	.99765	200	.99216
75	.99694	240	.99029
100	.99598	300	.98846
125	.99501		

%	π	%	π
	20°		
0	44.37	30	38.49
5	43.44	40	36.46
10	42.47	50	34.32
15	41.46	60.15	32.02
20	40.47		
25	39.56		

Grassi, 1851

%	t	π
0	13.4	47.7
58.67	15.5	26.0

Freyer, 1931

%	π	t	π
20°		30%	
6	44.31	10	37.42
16	41.29	20	36.80
30	36.80	30	36.38
45	31.67	40	36.14
		50	36.01
		60	35.93

Scott, Obenhaus and Wilson, 1934

%	π
35°	
4.4120	40.87
8.9594	40.04
16.415	38.44
34.901	34.22
58.262	27.74

Thomas and Perman, 1934

%	π(0-100atm)	%	π(0-100atm)
30°			
6.32	42.2	31.02	33.6
12.41	39.9	38.76	31.3
19.86	37.4	47.53	30.2

Viscosity and surface tension

Taylor and Ranken, 1903-04

N	η(water=1)	N	η(water=1)	N	η(water=1)
0°		15°		25°	
1	0.854	1	0.583	1	0.467
2	.778	2	.552	2	.458
3	.748	3	.544	3	.459

Getman, 1907

c	η	c	η
18°			
0	1058	45	959
10	1027	50	966
20	1000	60	985
25	986	70	1005
30	975	80	1046
35	974	90	1085
40	963	100	1164

c	η 10°	η 30°	η 50°
9.13	1240	780	560
19.90	1160	750	550
43.94	1110	750	570
67.12	1120	800	630
83.16	1230	860	670
100	1310	800	550

Grüneisen, 1905

M	η(water=1)	M	η (water=1)
18°			
0.05	0.99551	1	.9201
.10	.99079	5.245	.9778
.20	.98113	5.6	1.013
.50	.95509		

Herz, 1914 and 1917

N	η(water=1)	N	η(water=1)
	25°		
0	1	3.888	0.946
1.296	0.920	5.184	1.030
1.944	.914	6.200	1.193
2.592	.915		

Herz, 1917

N	η	N	η
	25°		
1.296	823	3.888	847
1.944	818	5.184	922
2.592	819	6.200	1068

Hrynakowski, 1927

sat.sol. η=860 50.50°

Tammann and Rabe, 1928

wt%	η			
	0°	10°	30°	75°
14.12	1540.1	1208.8	767.1	394.5
24.89	1417.1	1106.1	755.8	408.9
33.00	1354.9	1975.6	758.4	427.3
39.85	1304.7	1090.4	774.9	449.9
45.35	1343.0	1099.5	799.5	473.6

Herz and Hiebenthal, 1929

N	η	
	25°	70°
0.5	848.1	419.6
1.0	835.2	423.6
2.0	811.1	443.1
4.0	864.8	496.3

Traube, 1885

c	a^2	σ
	15°	
2.5	14.485	-
5	14.289	-
10	13.837	72.90
15	13.528	-
20	13.132	73.78

Sentis, 1897

mol%	t	σ	mol%	t	σ
0	25.1	72.5	3	17.5	74.7
0	13.5	74.0	5	20.2	75.1
1	18.8	73.8	6.64	20.0	76.0
1	13.2	74.5	10	19.7	78.5
2	18.0	74.1			

De Block, 1925

%	σ
	15°
0	73.26
14.8	73.63
24.5	73.91
46.7	75.75
56.0	77.21

Hrynakowski, 1927

sat.sol. σ=74.36 50.50°

Optical and electrical properties

Kremers, 1857

%	n_D
	17°
0	1.3332
37.93	.3960
54.96	.4405

Damien, 1881

%	n Hα	n Hβ	n Hγ
		20°	
0	1.33108	1.33706	1.34035
18.70	.35810	.36601	.37044
23.44	.37035	.37856	.38341
39.70	.39889	.41110	.41792
44.22	.40770	.42065	.42814
55.12	.43741	.45250	.46103

Wegner, 1889

%	n Hα	n D
	20°	
2.9845	1.335024	1.337003
5.0089	.337733	.339833
10.2725	.345182	.347274
15.1212	.352379	.354604
20.0390	.360361	.362785

Bender, 1890

N(15°)	t	$n_{H\alpha}$	t	n_D
0	20.0	1.33100	18.8	1.33297
1	20.4	.35164	21.2	.35372
2	21.3	.37156	22.0	.37416
3	20.5	.39115	20.9	.39419
4	20.6	.41057	21.1	.41403
5	20.6	.42973	21.5	.43360

N(15°)	t	$n_{H\beta}$	t	$n_{H\gamma}$
0	20.2	1.33703	20.3	1.34021
1	21.0	.35896	21.1	.36304
2	21.7	.38058	21.6	.38570
3	20.8	.40171	20.7	.40744
4	20.9	.42240	21.0	.42931
5	21.1	.44289	21.2	.45079

Jahn, 1891

c	n Hα	n D	n Hβ
		20°	
0	1.3395	1.3332	1.3375
17.535	.3555	.3555	.3608
33.407	.3753	.3753	.3817

Wagner, 1903

c	n_D	c	n_D
	17.5°		
0	1.33320	10.295	1.34650
0.299	.33358	.587	.34687
.598	.33397	.879	.34724
.896	.33435	11.170	.34761
1.193	.33474	.461	.34798
.489	.33513	.752	.34836
.785	.33551	12.043	.34873
2.080	.33590	.334	.34910
.375	.33628	.625	.34947
.669	.33667	.916	.34984
.963	.33705	13.206	.35021
3.288	.33743	.496	.35058
.553	.33781	.789	.35095
.848	.33820	14.076	.35132
4.142	.33858	.366	.35169
.436	.33896	.655	.35205
.730	.33934	.944	.35242
5.024	.33972	15.233	.35279
.318	.34010	.522	.35316
.612	.34048	.811	.35352
.906	.34086	16.100	.35388
6.199	.34124	.389	.35425
.492	.34162	.678	.35461
.785	.34199	.967	.35497
7.078	.34237	17.256	.35533
.371	.34273	.545	.35569
.664	.34313	.834	.35606
.957	.34350	18.123	.35642
8.250	.34388	.411	.35678
.543	.34426	.699	.35714
.835	.34463	.987	.35750
9.127	.34500	19.275	.35786
.419	.34537	.563	.35822
.711	.34575	.851	.35858
10.003	.34612	20.139	.35894

Chéneveau, 1907

%	n_D	%	n_D
	18.5°		
0	1.33315	25.48	1.3722
9.65	.3463	32.35	.3854
18.08	.3599	37.91	.3980
21.89	.3661		

Getman and Wilson, 1908

%	n_D	%	n_D
	20°		
0	1.33298	25	1.37145
5	.34008	30	.38115
10	.34665	35	.39125
15	.35460	40	.40220
20	.36250	50	.42835

Heydweiller, 1909

N	n_D
18°	
0.5	1.34411
1.0	.35484
2.0	.37581
4.0	.41599

Baxter, Boylston and al., 1911

%	n_D	%	n_D
	25°		
0	1.33246	3.1122	1.33658
0.4379	.33306	5.1029	.33927
0.9151	.33367	8.5864	.34422
1.2259	.33409	25.604	.37194
2.3607	.33554		

Lübben, 1913

N	%	n 3611.9Å	4419.9	4800.1	5086.0
			18°		
0	0	1.34754	1.33997	1.33766	1.33633
0.4922	7.81	.36053	.35148	.34880	.34726
0.9752	14.47	.37303	.36259	.35959	.35784
1.934	25.95	.39763	.38445	.38072	.37860
3.878	43.86	.44519	.42690	.42175	.41886

Grufki, 1913

N	%	n Hα	Hβ	Hγ
		18°		
0	0	1.33138	1.33732	1.34050
0.5015	7.96	.34204	.34871	.35240
1.015	15.06	.35286	.36032	.36449
2.009	26.97	.37326	.38218	.38731
4.015	45.41	.41272	.42444	.43143

Shibata and Hölemann, 1931

m	n_{He}	m	n_{He}	m	n_{He}
25°		35°		45°	
0	1.33270	0	1.33149	0	1.3300
1.0578	.35296	1.0766	.35287	1.0069	.3499
1.7670	.36686	1.09513	.36846	1.9969	.36742
2.3858	.37737	3.4145	.39164	3.5646	.39187
3.0811	.38839	4.8063	.41081	4.8280	.40907
4.7566	.41205				

Jahn, 1891

c	$(\alpha)_{magn.}$
20°	
0	1
17.535	2.0670
33.407	2.0441

Humburg, 1893

%	$(\alpha)^{mol}_{magn.}$
16°	
15.4670	18.9433
39.0736	18.9513

Schönrock, 1893

%	$(\alpha)_{magn.}$
16°	
0	1
32.4875	2.0592

Agerer, 1905

%	$(\alpha)_{magn.}$
18.5°	
10.491	2.175
22.784	.088
48.578	.062

Dewar and Fleming, 1897

t	ε	t	ε
	52.17%		
-199.0	2.72	-191.6	3.04
-196.0	.80	-170.6	7.67
-196.7	.87	-148.0	21.60

Okazaki, 1933

%	Verdet's constant (3441 Å)
28°	
7.52	0.05374
14.81	.06412
21.26	.07416
27.07	.08416
36.63	.10252
42.55	.11614

Okazaki, 1936

%	Verdet's constant (D)
	25°
0	0.013075
14.01	.01675
20.70	.01877
27.57	.02110
32.60	.02302
37.46	.02494
39.41	.02575
43.44	.02765
49.81	.03072

Kohlrausch, 1879

%	κ	%	κ
	18°		
5	347	40	3149
10	677	50	3991
20	1445	55	4199
30	2289		

Dennhardt, 1899

N	λ		
	0°	10°	18°
0.5	75.8	94.2	106.6
1	74.0	90.8	104.0
2	72.9	87.7	101.5
3	68.5	82.8	96.0

Taylor and Ranken, 1903-04

M	λ
0°	
1	70.0
2	69.5
3	68.0

Jones and Jones and Getman, 1904

N	λ	N	λ
		0°	
0.068	70.63	2.188	64.7
.137	70.0	2.735	64.0
.273	69.5	3.282	63.8
.547	67.4	3.829	63.6
1.094	67.0	4.545	61.2
.641	66.0		

Johnston, 1907

N	λ	N	λ
		100°	
0.0005	3980	0.25	3161
.0010	4054	0.5	2921
.0033	3784	1.0	2586
.0125	3730	2.0	2392
.06	3513	3.55	1973
.12	3353	4.29	1829

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	116.5	1.0	103.6
.1	114.9	2.0	99.5
.2	111.2	4.0	88.2
.5	106.2		

Sloan, 1910

N	λ	N	λ
	0°		
0.0470	75.60	1.50	70.43
.0753	74.55	1.81	70.81
.0942	74.13	2.41	70.43
.128	73.13	3.01	69.32
.188	72.36	3.61	67.65
.241	72.03	3.62	67.83
.301	71.56	4.22	65.20
.340	70.83	4.80	64.33
.602	70.65	5.42	59.17
.753	70.46	5.66	57.30
1.20	70.61		

Clausen, 1912

N	λ 6°	λ 18°	λ 30°
0.4987	81.66	106.3	132.2
0.9965	80.74	103.6	127.5
1.990	79.74	100.07	121.1
3.980	73.67	89.68	105.9

N	$\tau.10^4$ 6°	$\tau.10^4$ 18°	$\tau.10^4$ 30°
0.4987	407.2	530.1	659.5
0.9965	804.7	1032	1271
1.990	1587	1991	2410
3.980	2932	3569	4215

Thermal constants

Faasch, 1911

c	U
18°	
0.497	0.918
0.992	.840
1.997	.720
3.995	.546

c= moles KI in 1 litre water

Randall and Rossini, 1929

m	U	m	U
	25°		
0.00	0.9979	0.20	0.9611
.01	.9960	.35	.9357
.02	.9940	.50	.9119
.05	.9883	.75	.8753
.10	.9790	1.00	.8419

Scholz, 1892

N	Q diss. (cal/g)	N	Q diss. (cal/g)
0.0625	-35.24	1.0	-32.52
0.125	-35.46	2.0	-29.88
0.25	-34.82	4.0	-26.03
0.5	-33.88	7.5	-22.06

Wüst and Lange, 1923 and 1924

N	Q diss integral	N	Q diss integral
0.708	-4895	0.438	-4895
1.416	4772	1.063	4651
2.091	4666	1.76	4449
2.773	4557	2.39	4231
3.442	4455	3.11	4048
4.107	4359	3.79	3883
4.758	4268	4.43	3718
5.400	4183	5.08	3591
6.045	4100	5.72	3443
6.677	4023	6.28	3339
7.296	3949	6.99	3212
7.902	3882	7.61	3135
8.504	3817	8.21	3040
9.17	3757	8.80	2967
9.68	3700	9.40	2894
10.26	3645	9.98	2824
10.84	3593	10.55	2770
11.40	3544	11.13	2724
11.97	3499	11.68	2675
12.53	3455	12.24	2631

Jäger, 1891

%	heat conductivity coefficient
	at room temperature
0	100
20	96.8
40	77.8
60	65.1

Kapustinskii and Ruzavin, 1956

heat conductivity coefficient c.10^6

%	c	%	c
	25°		
5.75	1422	27.4	1285
10.4	1394	38.4	1205
17.7	1353	49.6	1102

Water + Potassium hexafluorophosphate (KPF_6)

Sarmonsakis and Low, 1955

%	f.t.	%	f.t.
3.560	0	18.22	55
4.242	4.5	20.29	60
5.270	10.5	22.49	65
6.400	16.2	24.61	70
7.30	20.2	27.29	75
8.35	25	29.38	80
9.69	30	31.96	85
11.15	35	34.03	90
12.95	40	35.88	95
14.82	45	38.30	100
16.48	50		

Water + Potassium nitride (KN_3)

Wohlgemuth, 1934 (fig.)

%	f.t.	%	f.t.
0	0	30	+10
10	-5	40	43
20	-10	50	100
26.2	-12.9 E		

Water (H_2O) + Potassium cyanide (KCN)

Corbet, 1926

41.7% f.t.=25°

Vasilieva, 1925

%	f.t.	%	f.t.
40.37	15.3	37.06	-14.1
40.14	9.0	36.33	-17.3
38.68	0	34.73	-28.1
37.55	-8.7	33.44	-29.8

Tammann and Rohmann,1929

P Kg	t	(Dλ/λ).100			
		0.001 N	0.01 N	0.1 N	0.5 N
500	20	5.71	4.19	3.62	3.31
	40	3.85	2.42	2.10	1.92
1000	20	9.42	6.73	5.94	5.32
	40	6.20	4.09	3.52	2.95
1500	20	12.1	8.45	7.21	6.50
	40	7.95	5.21	4.30	3.60
2000	20	14.1	9.53	7.62	6.61
	40	9.19	5.72	4.48	3.55
2500	20	15.2	9.63	7.43	6.30
	40	9.76	5.73	3.94	3.00
3000	20	15.5	9.17	6.50	5.21
	40	9.80	5.35	2.85	1.95

P Kg	t	(Dλ/λ).100		
		1.0 N	2.5 N	5.0 N
500	20	2.97	1.45	-0.09
	40	1.69	0.71	-0.04
1000	20	4.61	2.19	-0.22
	40	2.59	1.18	-0.20
1500	20	5.39	2.39	-0.92
	40	2.93	1.23	-
2000	20	5.51	2.10	-2.26
	40	2.29	0.49	-
2500	20	4.89	1.32	-3.84
	40	2.18	-0.38	-
3000	20	3.84	0.00	-5.51
	40	1.17	-1.53	-

P Kg	t	λ			
		0.001 N	0.01 N	0.1 N	0.5 N
1	20	131.8	127.1	116.5	108.1
	40	189.0	181.9	166.4	142.9

P Kg	t	λ		
		1.0 N	2.5 N	5.0 N
1	20	103.2	93.1	80.9
	40	134.2	120.5	101.8

Water (H_2O) + Potassium cyanate (KCNO)

Traube, 1895

%	d	%	d
	15°		
0	0.99913	17.286	1.10737
5.008	1.02951	23.584	.14633
9.748	.05883	33.147	.21254

Water (H_2O) + Potassium thiocyanate (KCNS)

Heterogeneous eqilibria

Tammann, 1885

t	p 0%	17.00%	31.07%	33.94%	44.92%
37.62	48.7	46.1	42.5	41.5	37.5
42.66	63.7	59.9	55.0	53.7	48.4
49.36	89.6	84.0	76.5	75.3	77.0
53.91	112.1	104.9	95.4	93.1	83.8
57.02	130.0	121.3	110.5	107.7	96.7
60.34	151.8	142.2	129.5	126.3	112.9
65.39	190.9	179.5	162.8	158.6	141.2
68.40	218.1	204.5	185.6	181.0	160.9
72.17	256.7	240.6	218.2	212.8	188.7
76.27	305.1	295.1	259.0	253.3	224.6
82.43	391.6	365.0	331.2	323.3	286.2
88.04	488.1	454.4	412.2	402.1	355.9
91.39	554.4	517.0	469.0	457.3	404.2
100.74	780.3	724.8	660.7	639.7	-

%	p	%	p
	100°		
6.98	742.8	53.63	479.4
10.54	730.8	57.91	444.5
20.93	691.1	58.45	438.7
31.96	635.2	61.46	411.7
43.08	563.2	72.34	407.9
46.10	543.1		

Gibson and Adams, 1933

%	p	%	p
	20.28°		
30.84	10.15	54.75	13.61
40.17	10.92	59.00	14,253
44.25	11.767	62.74	15.377
51.68	12.36		

Dingemans, 1939

t	p	t	p
	sat.sol.		
10	4.8	92	128.3
12	5.4	95	137.4
15	6.4	97	143.6
17	7.1	100	153.0
20	8.3	102	159.4
22	9.3	105	168.3
25	10.8	107	175.0
27	12.0	110	184.0
30	13.8	112	189.8
32	15.3	115	198.2
35	17.6	117	203.5
37	19.3	120	210.7
40	22.1	122	215.2
42	24.1	125	221.0
45	27.4	127	224.2
47	29.7	130	228.3
50	33.6	132	229.7
52	36.4	135	231.0
55	40.9	137	230.7
57	44.0	140	224.5
60	49.1	142	228
62	52.8	145	224
65	58.6	147	220
67	62.8	150	212
70	69.4	152	206
72	73.9	155	195
75	80.8	157	186
77	85.8	160	169
80	93.7	162	156
82	99.2	165	134
85	107.6	167	116
87	113.4	170	88
90	122.3	172	68

Hunter and Bliss, 1944

mol%	p	mol%	p
	30°		
4.42	30.24	10.95	27.74
4.47	30.15	21.50	23.76
10.90	27.83	21.73	23.83

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.926	1.2	0.892
0.2	0.911	1.4	0.892
0.3	0.904	1.6	0.893
0.4	0.900	1.8	0.894
0.5	0.897	2.0	0.898
0.6	0.896	2.5	0.903
0.7	0.895	3.0	0.908
0.8	0.894	3.5	0.912
0.9	0.894	4.0	0.917
1.0	0.893	4.5	0.921

Rüdorff, 1872

%	f.t.
4.76	-1.60
9.09	-3.25
13.04	-4.90
16.67	-6.50
20.00	-7.90
23.08	-9.55

Meyerhoffer, 1905

50.25% -31.2 E

Vasiliev, 1910

15.75 mol% -31.2° E

Kordes, 1926

mol%	f.t.
0	0
15.5	-31 E
100	+179

Chrétien and Hoffer, 1935

%	f.t.	%	f.t.
0	0	66	+6.7
20	-7	80	63
40	-22	97	142.3 tr.t. I + II
54.8	-30.1	100	177.2
60	-10		

E: -33.2 (1+2) (4+5)

Properties of phases

Traube, 1895

%	d	%	d
	15°		
0	0.99913	12.647	1.06379
3.481	1.01676	20.516	.10618
6.326	.03115	34.227	.18519

Heydweiller, 1909

N	%	d	N	%	d
		18°			
0.1	0.87	1.00346	1.0	8.37	1.0457
.2	1.72	.00826	2.0	15.97	.0915
.5	4.27	.02234	4.0	29.52	.1795

Schneider, 1910

N	d	N	d
	18°		
0	0.99862	1.005	1.0460
0.1005	1.00350	.975	.0905
.201	.00834	3.95	.1775
.5025	.02251		

Rubien, 1911

N	%	d	N	%	d
		18°			
0.0953	0.83	1.00341	0.9737	8.13	1.04492
.1916	1.64	.00809	1.9602	15.65	.09008
.1837	4.14	.02200	3.9513	29.05	.17742

Grufki, 1913

N	%	d	N	%	d
		18°			
0	0	0.99862	1.953	15.68	1.08942
0.4915	4.19	1.02119	3.941	29.12	.17649
0.9825	8.20	1.04483			

Gibson and Adams, 1933

%	d	%	d
	20.28°		
30.84	1.3757	54.75	1.2469
40.17	1.3481	59.00	1.2210
44.25	1.3177	62.74	1.1660
51.68	1.2963		

Gibson, 1935

%	d	%	d
	25°		
0.00	0.9970	38.95	1.2096
8.29	1.0380	44.37	.2448
14.82	.0716	54.97	.3168
26.49	.1356	64.54	.3870

Gibson, 1935

%	π (1-1000 bars)	%	π (1-1000 bars)
	25°		
0.00	39.35	38.95	27.48
8.29	36.62	44.37	26.03
14.82	34.52	54.97	23.35
26.49	31.13	64.54	21.14

Schneider, 1910

N	η (H_2O = 1)	N	η (H_2O = 1)
	18°		
0.1005	0.9975	1.005	0.9587
.201	.9915	1.975	.9500
.5025	.9768	3.95	1.0333

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0.1	1.33508	1.0	1.35069
.2	.33681	2.0	.36780
.5	.34204	4.0	.40018

Rubien, 1911

N	%	n_D	N	%	n_D
		18°			
0.953	0.83	1.33499	0.9737	8.13	1.35024
.1916	1.64	.33666	1.9602	15.65	.36722
.4837	4.14	.34176	3.9513	29.05	.39938

Grufki, 1913

N	%	n H_α	n H_β	n H_γ
			18°	
0	0	1.33140	1.33734	1.34054
0.4915	4.19	.33992	.34627	.34973
0.9825	8.20	.34822	.35502	.35877
1.955	15.68	.36452	.37216	.37646
3.941	29.12	.39632	.40562	.41101

Dhar, 1914

mol%	κ	mol%	κ
	30°		
2.14	2737	9.83	3293
4.91	3874	11.75	3425 sic
6.55	4152	12.95	3820

Heydweiller, 1909

N	%	λ	N	%	λ
		18°			
0.1	0.87	104.3	1.0	8.34	91.6
.2	1.72	102.0	2.0	15.97	86.8
.5	4.27	95.7	4.0	29.52	74.6

Heat properties

Urban, 1932

t	U 2.625 m	1.257 m	0.500 m
10	0.8078	0.8908	0.9501
15	.8120	.8935	.9517
20	.8146	.8950	.9524
25	.8160	.8958	.9530
30	.8162	.8963	.9534
35	.8156	.8963	.9537
40	.8142	.8963	.9537

Partington and Soper, 1929

mol%	Q diss(cal/mole KCNS) integral	differential	Q dil (by mole H_2O)
0	-5988	-5988	0.0
0.4	5795	5693	-0.51
0.9	5693	5494	1.99
1.4	5603	5325	4.17
1.9	5508	5157	7.02
2.3	5422	5003	10.475
2.9	5339	4859	14.40
3.4	5258	4727	15.585
3.8	5185	4514	22.84
4.2	5115	4492	28.035
4.7	5050	4380	33.50
5.2	4985	4270	-39.325
23.0	-3470	-2792	-203.4
23.6	3447	2762	212.35
24.2	3426	2736	220.8
24.8	3406	2716	227.7
25.3	3387	2699	233.9
25.9	3368	2685	233.9
26.4	3350	2674	239.05
27.0	3332	2667	246.05
27.5	3314	2661	248.14
28.0	3298	2656	250.38
28.5	3281	2651	352.00
29.0	3262	2647	354.20
29.5	3252	2644	355.36
30.0	3239	2641	357.14
30.1	3235	2640	-357.58

Hunter and Bliss, 1944

mol%	Q vap(cal/g H_2O)	Q dil(cal/mole H_2O)
	30°	
4.42	579.2	17.8
4.47	579.3	16.7
10.90	574.0	112.8
10.95	574.2	108.4
21.50	568.8	205.3
21.73	568.7	208.4

Water (H_2O) + Potassium hydroxide (KOH)

Heterogeneous equilibria

Wüllner, 1860

t	p 0%	5.54%	10.15%
11.70	10.255	9.756	9.158
12.10	10.562	10.063	9.465
13.95	11.863	11.265	10.966
15.15	12.813	12.165	11.517
15.30	12.947	12.299	11.552
16.35	13.825	12.928	12.230
19.40	16.785	15.790	14.795
20.25	17.685	16.641	15.549
21.82	19.423	18.380	17.237
23.65	21.659	20.464	19.169
25.53	24.261	22.768	21.375
26.98	26.500	24.710	23.220
27.93	27.985	26.395	24.705
28.60	29.101	27.713	25.522
30.65	32.750	30.861	28.852
32.13	35.600	33.612	31.524
34.95	41.827	39.641	37.155
35.70	43.536	41.450	39.164
36.64	45.940	43.159	40.477
37.75	48.500	45.770	42.790
40.10	54.300	51.369	48.095
40.16	54.400	51.421	48.345
42.82	63.650	60.072	56.505
45.32	72.486	68.561	63.499
45.65	73.900	69.831	65.561
47.28	79.995	75.682	70.983
49.80	90.983	86.058	80.587
50.90	96.012	93.652	86.153
53.38	108.870	103.421	96.885
55.43	119.740	113.403	107.383
57.57	133.562	126.823	119.250
59.95	148.451	140.651	132.251
62.63	167.762	158.813	149.127
64.91	186.070	176.590	166.360
67.00	204.376	193.903	182.633
70.08	234.000	221.962	209.935
72.05	254.634	242.488	229.847
72.50	259.570	245.770	-
74.90	287.328	272.035	-
77.18	316.200	299.800	-
80.20	357.571	339.171	-
82.30	387.568	368.029	-
85.38	439.890	417.560	-
87.28	473.704	449.484	-
90.48	535.130	507.278	-
92.77	583.220	553.890	-
95.30	640.830	608.090	-
97.38	692.040	657.122	-
99.20	738.500	702.908	-

t	p 14.05%	p 17.40%	p 20.03%
11.70	7.961	7.063	6.565
12.10	8.166	7.871	6.860
13.95	10.069	9.419	8.723
15.15	10.023	8.925	8.125
15.30	10.256	9.659	8.261
16.35	10.685	9.545	8.800
19.40	13.005	11.710	10.615
20.25	13.806	13.005	10.908
21.82	15.544	13.858	12.595
23.65	17.123	15.135	13.936
25.53	19.186	18.139	15.903
26.98	20.734	18.892	16.988
27.93	21.920	19.835	18.153
28.60	23.633	21.699	19.473
30.65	25.887	23.394	21.392
32.13	28.441	25.710	23.372
34.95	33.483	30.399	26.956
35.70	35.387	31.536	28.965
36.64	36.800	33.374	30.393
37.75	38.568	35.000	31.653
40.10	43.429	38.961	35.188
40.16	43.979	39.611	36.138
42.82	51.238	46.310	42.238
45.32	57.236	51.319	45.746
45.65	60.328	54.522	50.057
47.28	64.144	58.600	53.295
49.80	63.641	66.718	60.884
50.90	77.581	71.004	65.003
53.38	88.465	80.293	73.507
55.43	97.313	87.829	80.925
57.57	108.962	99.356	91.128
59.95	120.998	109.109	100.846
62.63	136.737	134.350	113.987
64.91	152.270	137.587	126.587
67.00	166.953	152.339	138.876
70.08	191.088	174.900	159.550
72.05	210.134	292.250	175.939
72.50	211.972	193.829	-
74.90	235.463	214.322	-
77.18	259.312	236.777	-
80.20	293.271	268.571	-
82.30	318.092	290.201	-
85.38	362.363	331.060	-
87.28	389.532	354.504	-
90.48	439.932	-	-
92.77	479.750	-	-
95.30	527.830	-	-
97.38	570.677	-	-
99.20	611.463	-	-

Tammann, 1885

%	p	%	p
100°			
2.41	746.9	17.54	629.5
4.69	734.2	22.65	568.9
5.88	726.0	29.97	465.8
7.93	710.8	33.66	402.5
9.03	704.5	37.18	353.5
11.67	683.5	38.35	373.0
12.80	675.9	42.64	272.7
14.76	657.2		

Dieterici, 1898

N	p	N	p
0°			
0	4.579	7.214	3.470
0.995	.489	11.94	2.593
1.939	.356	19.14	1.547
3.232	.196	31.23	0.638
4.846	3.898		

Smits, 1899

m	Dp	m	Dp
0°			
0.03035	0.00409	0.51342	0.07504
.03476	.00470	0.75044	.11170
.05564	.00763	1.0356	.15867
.0992	.01382	1.1912	.19505
.16626	.02321	2.5995	.48440
.33464	.04786	2.6422	.47601
.42374	.06454		

Boswell and Cantelo, 1922

N	Dp	N	Dp
23°			
0.930	0.028	9.400	0.611
1.800	.070	10.200	.643
3.600	.178	12.500	.813
5.300	.292	14.700	.882
7.100	.405		

Shibata, Odda and Furukawa, 1932

t·	p
sat.sol. (2+1)	
25	1.94
27.5	2.12
30	2.24
32.5	2.31

Galinker and Korobkov, 1951

mol%	P	mol%	P	mol%	P
300°		350°		400°	
42.5	5	46.0	5	51	5
36.8	10	41.0	10	46	10
29.5	20	31.0	30	42	15
23.8	30	25.5	45	35	30
18.8	40	20.5	60	26.5	60
10.5	60	14.3	80	21.5	80
8.5	65	8.5	100	14.5	110
4.5	75	5.2	125	10	140
1	85	1.8	150	7.5	155
				4.0	180
420°		460°			
57	5	57	10		
50	10	52	15		
42	25	47.5	35		
26.2	80	28	100		
23	100	24.3	130		
16	150	17.5	190		
10	190	14	220		

Korobkov and Galinker, 1956

mol%	P	mol%	P	mol%	P
300°		350°		374°	
49	5	50	8	52	11
37	15	38	23	40	29
28	27	28	47	29	53
21	40	22	64	23	73
16	48	16	86	17	100
11	64	11	110	11	133
5	75	6	136	8	150
		3	154	6	169
				3	199
400°		420°			
54	13	55	16		
42	32	43	38		
32	61	34	67		
24	89	26	102		
18	123	19	143		
14	150	16	168		
12	176	12	206		
8	212	10	239		
6	216				

Bolte, 1912

	t	p dissoc.
(1+2)	27.75	3.60
(1+1)	27.75	1.54
(1+1)	31.75	2.13
(4+3)	31.75	0.75

Gerlach, 1886

%	b.t.	%	b.t.
0	100	67.79	215
17.01	105	68.73	220
25.65	110	69.70	225
31.63	115	70.66	230
36.51	120	71.58	235
40.30	125	72.47	240
43.34	130	73.29	245
45.95	135	74.08	250
48.06	140	74.89	255
49.95	145	75.76	260
51.58	150	76.63	265
53.28	155	77.45	270
54.89	160	78.21	275
56.40	165	78.95	280
57.81	170	79.63	285
59.15	175	80.32	290
60.42	180	80.97	295
61.60	185	81.63	300
62.71	190	82.88	310
63.84	195	84.03	320
64.91	200	85.11	330
65.94	205	86.18	240
66.89	210		

Robinson and Stokes, 1949

m	osmotic coefficient
25°	
0.1	0.944
.2	.936
.3	.938
.4	.944
.5	.953
.6	.962
.7	.972
.8	.983
.9	.993
1.0	1.003
.2	.026
.4	.051
.6	.076
.8	.100
2.0	.125
2.5	.183
3.0	.248
3.5	.317
4.0	.387
4.5	.459
5.0	.524
5.5	.594
6.0	.661

Rüdorff, 1862

%	f.t.
2.31	-1.45
3.81	-2.55
6.28	-4.50
6.52	-4.75
8.55	-6.60

Pickering, 1893

%	f.t.	%	f.t.
	(1+1)		
57.84	41.80	64.573	105.0
58.516	49.0	64.61	105.80
58.83	53.70	66.415	110.0
60.29	69.75	66.72	118.60
60.423	70.0	68.54	127.90
62.46	90.00	70.22	134.65
62.499	88.5	71.84	139.15
	(2+1)		
45.81	-27	52.77	20.0
45.874	-23.2	53.388	22.5
48.12	- 9.	53.89	24.0
48.142	- 5.8	54.94	28.0
49.90	+ 3.5	54.986	27.5
50.467	8.8	55.81	30.0
51.37	12.5	56.716	32.8
51.852	17.0	56.82	32.5
	(4+1)		
34.352	-75	41.87	-34.2
37.327	-42.2	43.092	-32.7
40.095	-35.0	43.12	-33.0
40.56	-36.2	44.29	-33.0
	ice		
3.509	- 2.2	22.661	-30.5
6.839	- 4.6	27.029	-44.2
12.966	-11.2	30.833	-65.2
18.496	-20.7		

Jones and Getman, 1902, 1903 and 1904

N	f.t.	N	f.t.
0.2	-0.719	1.5	-5.964
0.5	-1.805	2.0	-8.485
1.0	-3.890	2.5	-11.302

Jones, 1904; Jones and Getman, 1904; Jones and Bassett, 1905

N	f.t.	N	f.t.
0.05	-0.183	2.00	- 8.42
.10	-0.357	3.00	-14.0
.20	-0.710	4.00	-22.5
.30	-1.050	5.00	-32.5
.40	-1.390	6.00	-43.5
.60	-2.143	7.00	-57.5
.80	-2.885	8.00	-76.0
1.00	-3.773		

Cohen-Adad and Michaud, 1956 (fig.)

%	f.t.	E	tr.t.	
0	0	-	-	
10	-6	-62.8	-	
20	-25	-62.8	-65.7	
30.9	-62.8	-62.8	-65.7	
39	-40	-62.8	-65.7	
43	-30	-62.8	-65.7	
45.2	-34	-34	-	(4+1)
50	+10	-34	-	
58.1	33	-34	+33	
70	135	-	33	
76	145	-	-	(1+1)
80	140	+99	85	
87	99	99	85	
90	170	99	85	
95	310	99	85	and 242
100	401	-	-	

(4+1) f.t.= -33.7° (congruent)

Properties of phases

Frankenheim, 1847

t	d	t	d
0	1.2737	55	1.2431
5	.2711	60	.2402
10	.2686	65	.2372
15	.2656	70	.2345
20	.2630	75	.2319
25	.2601	80	.2290
30	.2573	85	.2260
35	.2546	90	.2233
40	.2518	95	.2202
45	.2490	100	.2162
50	.2461		

Wiedemann, 1856

%	d	%	d
		16.6°	
0	0.9989	11.919	1.1125
4.365	1.0399	15.463	1.1466
8.284	1.0770		

Schiff, 1858

%	d	%	d	%	d
		15°			
1	1.008	25	1.229	48	1.510
2	1:016	26	1.240	49	1.526
3	1.024	27	1.251	50	1.538
4	1.032	28	1.263	51	1.551
5	1.040	29	1.277	52	1.564
6	1.048	30	1.287	53	1.577
7	1.057	31	1.299	54	1.589
8	1.064	32	1.310	55	1.603
9	1.073	33	1.323	56	1.617
10	1.082	34	1.335	57	1.629
11	1.091	35	1.348	58	1.640
12	1.109	36	1.360	59	1.654
13	1.110	37	1.373	60	1.666
14	1.118	38	1.386	61	1.681
15	1.127	39	1.399	62	1.694
16	1.136	40	1.410	63	1.704
17	1.145	41	1.424	64	1.716
18	1.154	42	1.437	65	1.727
19	1.165	43	1.449	66	1.738
20	1.176	44	1.461	67	1.749
21	1.187	45	1.474	68	1.766
22	1.197	46	1.487	69	1.778
23	1.208	47	1.498	70	1.788
24	1.219				

Berthelot, 1873

Density at 10 - 16°

Hager, 1876

%	d	%	d
	17.5°		
1.19	1.009	28.60	1.280
2.38	1.020	29.79	1.292
3.57	1.030	30.98	1.305
4.77	1.041	32.17	1.318
5.96	1.053	33.36	1.330
7.15	1.064	34.55	1.343
8.34	1.075	35.74	1.356
9.53	1.087	36.93	1.370
10.79	1.098	38.12	1.384
11.91	1:110	39.31	1.398
13.10	1.122	40.51	1.412
14.29	1.134	41.70	1.426
15.48	1.146	42.89	1.440
16.68	1.157	44.08	1.454
17.87	1.169	45.27	1.468
19.06	1.181	46.46	1.482
20.25	1.198	47.65	1.498
21.44	1.205	48.84	1.512
22.63	1.217	50.03	1.528
23.83	1.229	51.22	1.543
25.02	1.242	52.42	1.558
26.21	1.255	53.61	1.574
27.40	1.267		

Tännermann, 1885

%	d	%	d
		15°	
0	0.9991	23.586	1.2111
5.958	1.0469	28.304	.2637
12.804	.1049	33.695	.3288
18.195	.1558		

Kohlrausch, 1879

%	d	%	d
		15°	
4.2	1.0382	25.2	1.2439
8.4	.0776	29.4	.2880
12.6	.1177	33.6	.3332
16.8	.1588	37.0	.3803
21.0	.2008	42.0	.4298

Traube, 1885

g	d	g	d
		15°	
5	1.0382	15	1.1064
10	1.0731	20	1.1369

g = g KOH in 100 cc water .

Pickering, 1894

%	d	%	d	%	d
		15°			
0	0.99918	18	1.16875	36	1.35485
1	1.00834	19	.17855	37	.36586
2	.01752	20	.18839	38	.37686
3	.02671	21	.19837	39	.38793
4	.03593	22	.20834	40	.39906
5	.04517	23	.21838	41	.41025
6	.05443	24	.22849	42	.42150
7	.06371	25	.23866	43	.43289
8	.07302	26	.24888	44	.44429
9	.08240	27	.25918	45	.45577
10	.09183	28	.26954	46	.46733
11	.10127	29	.27997	47	.47896
12	.11076	30	.29046	48	.49067
13	.12031	31	.30102	49	.50245
14	.12991	32	.31166	50	.51430
15	.13955	33	.32236	51	.52622
16	.14925	34	.33313	52	.53822
17	.15898	35	.34396		

Moore, 1895 - 1896

N	d	N	d
		18°	
0.00	0.9987	1.82	1.0864
0.456	1.0212	4.00	1.1793
0.92	1.0433	6.8	1.2900

Jones, 1904; Jones and Getman, 1904; Jones and Bassett, 1905

N	d	N	d
	0°		
0.05	1.0030	2.00	1.0926
.10	.0050	3.00	.1340
.20	.0100	4.00	.1761
.30	.0150	5.00	.2144
.40	.0197	6.00	.2507
.60	.0295	7.00	.2873
.80	.0388	8.00	.3323
1.00	.0482		

Zecchini, 1905

%	t	d	%	t	d
16.6210	72	1.16174	4.7880	23.4	1.04865
22.7770	70	.22346	11.9853	23.6	.12958
3.0530	23.2	.03368	12.0278	22.8	.13101
3.2423	23.4	.03369	19.5946	24.7	.21514
4.5924	24.1	.04836	28.8692	22.4	.31853
			34.7400	21.6	.37629

Chéneveau, 1907

%	d	%	d
	18°		
0	0.9986	9.22	1.0845
3.23	1.0286	11.99	.1108
6.30	.0566	14.63	.1358
7.78	.0708		

Carstens, 1924

%	d	%	d
	18°		
0	0.9986	25.24	1.2405
3.71	1.0330	30.74	.2980
7.72	.0697	37.77	.3750
10.21	.0930	44.70	.4470
14.90	.1380	51.67	.5300
20.90	.1970		

Faust, 1927

N	%	d
	20°	
0.257	1.43	1.013
1.363	7.02	.064
2.813	13.99	.1303
6.434	28.38	.2743

Hitchcock and Ilkenny, 1935

N	d	N	d	N	d
20°		30°		40°	
0.8385	1.0385	3.522	1.1536	0.8316	1.0313
1.6396	.0748	7.157	.2923	1.6273	.0668
2.495	.1119	10.210	.3969	2.475	.1032
3.246	.1435			3.220	.1344
3.964	.1731			3.932	.1636
4.879	.2097			4.840	.1999
5.606	.2374			5.561	.2279
6.366	.2669			6.315	.2568

Åkerlöf and Bender, 1941

%	N	d 0°	d 10°	d 20°	d 30°
0	0.0000	0.99982	0.99948	0.99800	0.99553
2	.3637	1.01927	1.01822	1.01623	1.01341
4	.7427	.03849	.03681	.03437	.03124
6	1.1377	.05773	.05548	.05262	.04922
8	.5499	.07706	.07429	.07106	.06739
10	.9804	.09650	.09327	.08970	.08579
12	2.4305	.11607	.11243	.10855	.10433
14	.9016	.13576	.13177	.12763	.12332
16	3.3951	.15557	.15130	.14694	.14245
18	.9126	.17546	.17098	.16645	.16183
20	4.4560	.19542	.19082	.18617	.18145
22	5.0273	.21532	.21069	.20600	.20122
24	.6287	.23575	.23094	.22614	.22132
26	6.2625	.25642	.25145	.24650	.24154
28	.9316	.27730	.27219	.26712	.26204
30	7.6389	.29840	.29318	.28799	.28280
32	8.3878	.31973	.31442	.30913	.30384
34	9.1821	.34132	.33591	.33053	.32514
36	10.0260	.36314	.35766	.35219	.34672
38	10.9244	.38524	.37969	.37413	.36857
40	11.8826	.40765	.40199	.39634	.39071
42	12.9070	.43037	.42459	.41884	.41313
44	14.0046	.45347	.44751	.44164	.43585
46	15.1834	.47700	.47079	.46477	.45889
48	16.4529	.50097	.49442	.48820	.48222
50	17.8240	.52565	.51861	.51211	.50600

%	d 40°	d 50°	d 60°	d 70°
0	0.99219	0.98809	0.98330	0.97789
2	1.00985	1.00562	1.00076	.99529
4	.02749	.02315	.01822	1.01269
6	.04528	.04083	.03583	.03026
8	.06329	.05872	.05366	.04804
10	.08153	.07686	.07174	.06607
12	.10002	.09526	.09008	.08437
14	.11877	.11392	.10868	.10293
16	.13779	.13285	.12756	.12176
18	.15706	.15205	.14670	.14087
20	.17659	.17151	.16610	.16024
22	.19630	.19116	.18750	.17981
24	.21635	.21115	.20565	.19973
26	.23651	.23128	.22576	.21982
28	.25691	.25159	.24602	.24006
30	.27758	.27127	.26653	.26054
32	.29852	.29302	.28733	.28129
34	.31974	.31416	.30840	.30232
36	.34124	.33558	.32976	.32364
38	.36303	.35729	.35141	.34524
40	.38510	.37930	.37335	.36713
42	.40747	.40160	.39560	.38932
44	.43015	.42422	.41817	.41182
46	.45315	.44718	.46108	.43465
48	.47645	.47045	.46431	.45779
50	.50022	.49419	.48802	.48139

Kurochkin, 1929

g	Dv
2.02	0.52
5.74	3.0
19.94	9.8

g= g KOH in 100 cc water

Schmidt, 1859

%	t	π
12.1	19.1	34.7
19	15.7	28.6
24.3	12.5	24.6
28.8	19.7	21.6

Carstens, 1924

%	π	%	π
18°			
0	49.1	25.24	23.2
3.71	43.7	30.74	20.2
7.72	38.7	37.77	16.4
10.21	36.0	44.70	14.4
14.90	31.4	51.67	10.7
20.90	25.9		

Wiedemann, 1856

%	η
16.6°	
0	1095
4.365	1225
8.284	1393
11.919	1572
15.463	1785

Moore, 1895-6

N	η	N	η
18°			
0.00	1058	1.00	1174
.25	1090	1.82	1282
.456	1106	2.00	1308
.5	1111	4.00	1670
.92	1162	6.80	2458

Faust, 1927

N	%	η
20°		
0.257	1.43	1034
1.363	7.02	1150
2.813	13.99	1382
6.434	28.38	2454

Hitchcock and Ilkenny, 1935

N	η	N	η	N	η
20°		30°		40°	
0.8385	1091.3	3.522	1223	0.8316	728.8
1.6396	1191.1	7.157	2030	1.6273	800.6
2.495	1311.4	10.210	3263	2.475	890.5
3.246	1438.3			3.220	978.8
3.964	1577.4			3.932	1071.7
4.879	1781.3			4.840	1205.4
5.606	1977.6			5.561	1328.5
6.366	2200.8			6.315	1471.6

Traube, 1885

g	σ
15°	
5	76.84
10	75.38
15	77.13
20	78.67

g= g KOH in 100 cc water

Faust, 1927

N	%	σ
	20°	
0.257	1.43	73.0
1.363	7.02	74.9
2.813	13.99	77.9
6.434	28.38	87.5

Zecchini, 1905

%	t	n_D	%	t	n_D
16.6210	72	1.36714	4.7880	23.4	1.34536
22.7770	70	.37895	11.9853	23.6	.36418
3.0530	23.2	.34210	12.0278	22.8	.36425
3.2423	23.4	.34169	19.5946	24.7	.38247
4.5924	24.1	.34505	28.8692	22.4	.40032
			34.7400	21.6	.41334

Briner, 1906

%	n_D	%	n_D
	18°		
0	1.3331	56.57	1.3560
33.33	.3425	60.00	.3590
42.86	.3469	62.88	.3618
47.81	.3439	63.64	.3627
50.00	.3511	66.67	.3663
55.55	.3552	68.05	.3685

Chéneveau, 1907

%	n_D	%	n_D
	18°		
0	1.3332	9.22	1.3509
3.23	.3394	11.99	.3561
6.30	.3451	14.63	.3612
7.78	.3482		

Fleming and Dewar, 1897

t	ε	t	ε	t	ε
50%		20%		10%	
-203.2	15.6	-201.5	41.2	-203.4	63.0
-149.8	22.9	-198.7	45.6	-202.5	67.2
-135.6	27.8	-180.0	46.4	-200.8	75.2
		-162.3	48.9	-197.2	87.5
				-179.0	89.5
				-170.9	90.8
5%		2.5%		-159.6	91.5
				-150.2	92.7
-204.8	123	-201.0	20.8	-141.1	95.0
-200.7	126	-166.0	29.0	-129.5	99.0
-196.2	131	-137.2	45.5	-117.1	272 ?
-181.5	135				
-170.1	133	0.05%			
-158.7	131	-203.4	10.7	-158.2	18.8
-147.3	132	-201.5	11.9	-152.0	20.0
-135.9	134	-196.0	14.7	-141.2	23.9
-127.7	143	-179.0	16.2	-132.1	35.1
		-169.5	17.1	-120.3	50.2 ?

Okazaki, 1933

%	Verdet's constant (3441 Å)
28°	
5.63	0.04618
9.94	.04805
16.57	.05050
19.91	.05233
27.06	.05542
33.35	.05798
40.75	.06126
43.03	.06218

Kohlrausch, 1879

%	κ	%	κ
	18°		
4.2	1458	25.2	5380
8.4	2712	29.4	5411
12.6	3748	33.6	5198
16.8	4540	37.0	4767
21.0	5086	42.0	4193

Jones and Getman, 1902, 1903 and 1904

N	λ	N	λ
	0°		
0.06	125.30	1	113.33
.12	122.80	1.5	106.35
.50	118.40	2	101.15
		2.5	97.40

Jones, 1904; Jones and Getman, 1904; Jones and Bassett, 1905

M	λ	N	λ
	0°		
0.05	136.0	2.00	102.31
.10	132.5	3.00	88.85
.20	131.0	4.00	77.60
.30	128.0	5.00	65.74
.40	126.0	6.00	56.19
.60	123.0	7.00	48.65
.80	118.0	8.00	38.77
1.00	116.94	9.00	32.25

Jaquerod, 1909

t	κ			
	56.0 g/1	94.2 g/1	140.8 g/1	213.5 g/1
18	1832	2822	3785	4761
20	1902	2937	3926	4957
25	2073	3200	4289	5447
30	2244	3462	4652	5937
35	2416	3724	5015	6427
40	2588	3996	5378	6917
45	2760	4260	5741	7407
50	2932	4525	6104	7898
55	3104	4790	6467	8488
60	3276	5050	6830	8879
65	3448	5310	7193	9369
70	3620	5570	7556	9860
75	3793	5830	7919	10350
80	3965	6090	8282	10841

g/1	κ			
	18°	25°	50°	75°
25	867	975	1470	1690
50	1156	1872	2655	3380
56.2	1837	2080	2940	3800
75	2348	2662	3780	4800
100	2956	3354	4740	6100
112.3	3224	3667	5190	6725
125	3484	3957	5605	7278
150	3935	4462	6390	8275
168.5	4207	4785	6835	8926
175	4292	4892	6995	9143
200	4604	5252	7590	9940
224.6	4882	5567	8145	10685

t	κ	t	κ	t	κ	t	κ
56.0 g/1		94.2 g/1		140.8 g/1		213.5g/1	
19.1	1870	20.7	2968	21.6	4044	23.4	5301
52.4	3021	50.6	4563	51.3	6202	50.5	7958
80.1	3962	80.2	6098	80.2	8291	80.1	10861

Thermal constants

Berthelot, 1873

Q diss at 10-16°

Hammerl, 1879

%	U	%	U
	12-30°		
7.48	0.900	21.42	0.807
9.27	.833	20.96	.780
11.62	.859	30.35	.737
13.29	.845	38.97	.697

Jaquerod, 1909

c	U	c	U
	18°		
2.5	0.960	17.5	0.820
5.0	.932	20.0	.805
7.5	.906	22.5	.790
10.0	.880	25.0	.775
12.5	.857	27.5	.759
15.0	.838	30.0	.744

Richards and Rowe, 1921

mol%	U	mol%	U
	20°		
3.9	0.8614	0.5	0.9773
2	.9211	0.25	.9882
1	.9556		

mol% initial	mol% final	Q dil
	20°	
9.9	3.9	+334
-	2	366
-	1	371
-	0.5	388
-	9.25	408

Jäger, 1891

%	thermal conductivity coefficient
0	100
21	95.5
42	90.6

XLVIII . WATER + POTASSIUM OXYSALTS .

Water (H_2O) + Potassium nitrite (KNO_2)

Tammann, 1815

%	p	%	p
100°			
12.30	723.1	52.39	518.9
18.51	699.9	55.64	494.4
26.16	666.8	62.37	448.3
34.86	623.8	64.40	433.9
40.52	592.7	69.61	389.5
44.01	572.3		

Oswald, 1914

Sat.sol. b.t.= 132°/758.5 mm

%	f.t.	%	f.t.
16.1	-4.1	75.75	+25
24.1	-7.6	77.0	40
40.2	-13.8	77.5	55.5
50.1	-18.6	78.6	75
61.7	-24.6	80.5	100
69.8	-30.0	80.7	111
71.8	-31.6 E	81.15	119
73.2	-6.5	81.8	125
73.6	0	100	297.5

Kordes, 1926

mol %	f.t.
0	0
35	-32 E
100	297

Bureau, 1935

%	f.t.	%	f.t.
0	0	60	-35
10	-4	64	-40
20	-8	70	-17
30	-13	73	-9
40	-20	78	+60
50	-26.5		

Bureau, 1935 and 1937

%	f.t.	E	tr.t.
13.01	-5.20	-40.0	-
27.85	-12.25	-40.2	-
41.0	-20.45	-40.3	-
58.8	-34.40	-40.2	-
64.9	-40.2	-40.2	- (1+2)
68.4	-26.8	-40.2	-
71.9	-8.9	-40.2	-8.9
80.0	-	-40.45	-8.55
88.5	-	-40.2	-8.9
92.0	-	-	-8.8

%	f.t.	d
74.3	20.0	1.647
76.36	56.0	.656
77.0	64.7	.655
78.5	79.5	.659
78.9	98.5	.670

Oswald, 1914

%	d
17.5°	
20.7	1.1336
36.3	.2540
48.3	.3341
64.6	.5363
74.5 (satd)	.6462

Rehbinder, 1926

%	d	%	d
20°			
75.4	1.635	40.5	1.279
71.9	.596	30.0	.192
65.8	.536	20.0	.111
60.1	.467	0.0	0.9982
50.0	.361		

%	σ	%	σ
20°			
75.4	95.0	40.5	79.5
71.9	91.7	30.0	77.4
65.8	88.0	20.0	76.0
60.1	85.0	0.0	72.75
50.0	81.7		

Water (H_2O) + Potassium chlorate ($KClO_3$)

Tammann, 1885

%	p	%	p
	100°		
3.77	753.0	23.47	708.4
9.36	741.9	27.37	696.9
13.67	732.0	33.59	677.6
16.44	724.9		

t	p		
	0%	10.24%	14.14%
64.50	183.5	180.9	-
66.83	203.5	200.2	200.0
70.95	243.6	238.6	-
73.82	275.3	270.2	270.0
75.74	344.3	292.9	292.4
79.21	377.5	337.5	337.1
81.49	432.3	368.3	367.3
84.91	488.2	421.6	420.1
88.05	553.9	476.3	474.8
91.37	605.9	540.6	538.3
93.77	677.1	590.3	588.4
96.80	771.6	660.7	657.1
100.43	275.3	751.4	747.8

Legrand, 1835

%	b.t.
12.77	101
22.65	102
30.51	103
36.93	104
38.08 sat.sol.	104.2

Kremers, 1856

Sat. sol. b.t. = 105°

Gerlach, 1886

%	b.t.	%	b.t.
0	100	26.36	102.5
6.10	100.5	30.84	103
11.66	101	34.81	103.5
16.81	101.5	38.35	104
21.76	102	40.90	104.4

Buchanan, 1899

%	b.t.	%	b.t.
	768.2 mm		
0	100.30	31.86	103.30
18.07	101.81	35.45	103.75
23.27	102.30	37.74	104.01
27.78	102.80	37.94	104.08
	741.4 mm		
0.00	99.31	13.85	100.31
4.60	99.73	15.95	100.62
6.28	99.83	18.17	100.84
8.09	99.92	20.07	101.02
0.29	100.01	23.19	101.32
10.62	100.11	27.44	101.83
11.78	100.21	31.70	102.33
12.81	100.31		
	727.5 mm		
0	98.78	26.82	101.14
17.66	100.20	29.50	101.44
19.80	100.42	32.04	101.74
21.96	100.63	34.52	102.05
23.96	100.83	36.73	102.35
	620.8 mm		
0	94.43	23.77	96.43
14.09	95.52	27.45	96.84
16.31	95.72	30.91	97.24
18.67	95.95	34.28	97.65
20.72	96.13	34.04	97.73
	550.35 mm		
0.00	91.20	29.22	93.73
19.25	92.72	32.66	94.13
24.51	93.23	32.68	94.24

Kahlenberg, 1914

%	b.t.	%	b.t.
	760 mm		
3.61	100.35	22.89	102.14
7.51	100.66	26.16	102.54
11.37	101.03	30.05	103.04
14.61	101.34	32.85	103.49
19.01	101.75		

Gérardin, 1865

%	f.t.	%	f.t.
8.67	28	15.47	47
10.95	35	22.54	65
12.58	40		

Tilden and Shenstone, 1884

%	f.t.	%	f.t.
42.43	120	59.68	160
49.72	136	64.67	190

Etard, 1894

%	f.t.	%	f.t.
2.6	-0.5	31.2	92
2.4	-0.3	37.2	106
3.5 ?	+4.5	47.0	130
2.9	4.5	59.8	171
4.7	11	62.1	180
6.1	19	63.1	190
8.9	29	64.2	200
9.9	36	66.0	207
11.4	42	87.0	300
15.1	56	96.7	330
16.6	58	96.0	358

Pawlewski, 1899

%	f.t.	%	f.t.
3.06	0	16.85	55
3.67	5	18.97	60
4.27	10	20.32	65
5.11	15	22.55	70
6.76	20	24.82	75
7.56	25	26.47	80
8.46	30	29.25	85
10.29	35	34.36	90
11.75	40	33.76	95
13.16	45	35.83	100
15.18	50		

Carlson, 1910

%	f.t.	%	f.t.
3.19	0	27.38	80
6.89	20	36.11	100
12.12	40	37.46(std)	104 b.t.
19.35	60		

Calzolari, 1912

%	f.t.	%	f.t.
4.24	8	9.31	30
6.67	19.8	36.43	99

Tschugaeff and Chlopin, 1914

%	f.t.	%	f.t.
17.37	53	28.53	81
23.25	68	30.46	88

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
65.1	177	83.5	242
70.7	195	90.7	277
73.1	203	92.2	284
75.8	212	95.7	305
78.8	222	100	370

Schmidt, 1859

d	π
17.0°	
1.024	40.9
.019	42.9
.009	45.4

Bindel, 1890

%	d	%	d
20°			
6.37	1.0418	14.54	1.0970
10.19	.0676	18.49	.1250
11.98	.0795	25.39	.1777

Carlson, 1910

%	t	d	%	t	d
3.19	0	1.021	27.38	80	1.165
6.89	20	.045	36.11	100	.219
12.12	40	.073	37.46	104	.230
19.35	60	.115			

Hrynakowski, 1927

Sat.sol.	50.40°	d=1.1991 η=722 σ=81.86

Bindel, 1890

%	U	%	U
20°			
6.37	0.915	14.54	0.819
10.19	.869	18.49	.778
11.98	.847	25.39	.715

%	Q diss (by mole chlorate)
20°	
6.37	-6820
10.19	-6457
11.98	-6417
14.54	-6283
18.49	-6166
25.39	-6008

Water (H_2O) + Potassium bromate ($KBrO_3$)

Tammann, 1885

%	p
100°	
6.18	751.6
11.76	742.8
17.50	730.5
28.01	708.1

Kremers, 1854

%	f.t.	%	f.t.
3.24	0	19.23	60
5.14	10	27.10	80
6.60	20	33.89	100
12.15	40		

Kremers, 1856

%	f.t.	%	f.t.
3.02	0	18.55	60
6.48	20	25.32	80
11.69	40	33.22	100

sat.sol. b.t.= 104°

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
43.6	134	72.6	230
48.4	149	72.6	249
51.1	160	74.4	254
53.5	167	77.2	265
54.1	170	79.1	274
57.3	172	81.1	279
59.9	186	81.4	286
63.2	193	83.1	297
64.2	204	86.4	312
67.4	211	100	382
70.6	226		

Ricci, 1934

%	f.t.	d	%	f.t.	d
3.642	5	1.024	8.785	30	1.062
4.510	10	.035	10.13	35	.074
5.397	15	.042	11.58	40	.083
6.460	20	.048	13.08	45	
7.533	25	.054	14.69	50	

Water (H_2O) + Potassium iodate(KIO_3)

Rosenheim and Liebknecht, 1899

%	b.t.	%	b.t.
0	100.000	6.72	100.277
2.74	100.112	9.73	100.409
3.49	100.137	10.90	100.448
5.98	100.249	18.55	100.791

Kremers, 1856

%	f.t.	%	f.t.
4.523	0	15.63	60
7.53	20	19.92	80
11.42	40	24.39	100

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
26.1	117	46.8	231
27.4	126	48.6	243
31.4	147	50.5	253
34.1	160	51.6	265
37.4	177	53.1	269
41.6	201	56.5	291
42.6	206	58.0	300
44.1	220	100	560

Randall and Taylor, 1941

m	d	m	d
25°			
0.06279	1.0082	0.2846	1.0468
.1985	.0320	.3989	.0669
.2387	.0387	.40645	.0680

m	U	m	U
25°			
0.0434	0.9968	0.1840	0.9962
.0521	.9967	.2814	.9970
.0628	.9966	.2845	.9973
.0973	.9962	.3774	.9985
.1089	.9963		

Water (H_2O) + Potassium nitrate (KNO_3)
Heterogeneous equilibria

Wüllner, 1858

t	p 0%	4.76%	9.09%	13.04%
38.4	50.38	49.89	49.39	48.90
42.4	62.33	61.64	61.04	65.25
46.7	78.07	77.18	76.29	75.40
47.0	79.09	78.10	77.31	76.71
50.7	95.29	94.01	92.82	91.53
53.1	107.16	105.88	104.69	103.40
53.9	111.41	110.13	109.15	107.56
60.9	155.13	153.25	151.57	149.91
61.3	158.62	156.94	155.26	153.28
61.6	160.06	158.09	156.11	153.93
61.7	160.94	-	157.58	-
65.2	188.64	-	184.10	-
65.8	193.74	191.62	189.25	186.83
69.6	229.13	-	223.60	-
70.3	234.18	231.42	228.76	226.13
70.6	239.45	236.79	234.32	230.37
72.3	257.37	254.51	252.04	248.59
72.7	263.77	260.42	258.06	253.81
75.9	299.60	296.15	293.39	289.05
76.3	304.64	300.90	297.93	294.29
78.6	334.97	-	327.59	-
81.0	369.28	364.85	361.40	356.18
81.6	378.36	374.13	368.91	364.76
82.6	393.72	388.80	383.87	378.94
83.0	400.10	-	390.45	-
86.3	455.68	450.37	445.44	439.73
87.9	484.81	-	473.99	-
91.4	554.09	-	542.78	-
93.3	595.09	588.21	581.32	574.64
94.2	615.29	-	600.34	-
95.3	640.83	-	626.09	618.22
97.0	682.03	-	665.52	-
100.6	776.48	765.19	755.27	744.27
100.8	782.04	-	760.42	-

t	p 16.67%	20.0%	23.08%
38.4	48.50	48.00	-
42.4	59.66	59.06	-
46.7	74.70	74.11	-
47.0	75.72	74.63	-
50.7	90.84	-	-
53.1	102.70	101.31	-
53.9	107.17	105.78	-
60.9	148.61	146.82	-
61.3	152.19	150.51	-
61.6	152.94	-	-
61.7	153.58	-	149.97
65.2	179.68	-	175.99
65.8	185.05	182.58	-
69.6	218.86	-	213.43
70.3	224.16	221.33	-
70.6	-	-	-
72.3	246.12	243.46	-
72.7	251.63	248.20	-
75.9	286.63	282.69	-
76.3	291.34	287.78	-
78.6	-	-	312.12
81.0	353.28	349.68	-
81.6	362.40	358.17	-
82.6	-	-	-
83.0	-	-	371.33
86.3	436.48	431.52	-
87.9	-	-	451.05
91.4	-	-	516.61
93.3	567.36	560.37	-
94.2	-	-	566.98
95.3	611.42	-	-
97.0	-	-	633.78
100.6	736.02	-	-

Nicol, 1884

t	p unknown %	sat.sol.
65	152.9	-
75	231.5	221.0
85	341.7	314.7
95	499.3	427.1

Tammann, 1885

%	p	%	p	%	p
		100°			
3.88	751.8	20.03	712.0	55.11	590.5
5.46	748.2	25.31	695.8	58.13	575.6
5.73	747.9	27.38	689.2	64.15	544.2
7.06	743.9	40.03	647.9	67.36	522.1
10.52	735.5	43.82	636.3	68.36	515.4
14.80	724.9	49.14	616.1		

t	p 0%	11.37%	29.12%	46.33%	53.63%
64.92	187.0	181.0	174.9	160.0	151.1
67.55	210.1	202.8	195.7	179.4	168.4
69.77	231.5	223.4	215.4	197.4	185.2
72.94	265.3	255.9	247.1	225.7	212.1
75.00	289.3	279.7	270.1	246.0	241.6
77.62	322.6	310.9	300.2	273.2	256.7
80.75	366.4	354.8	342.5	310.6	291.5
83.84	414.5	400.3	386.8	351.1	329.2
89.55	517.1	502.0	487.1	447.1	423.0
90.67	539.5	521.3	504.0	460.0	433.6
91.55	557.6	537.6	518.5	469.3	440.0
95.32	641.5	617.2	597.6	540.0	506.2
98.64	723.8	-	667.1	604.3	567.0
99.96	759.0	-	705.3	636.3	596.1

t	p 0%	15.05%	t	p 0%	15.05%
55.83	122.9	118.7	81.35	375.4	361.5
57.64	133.9	128.8	88.38	494.5	476.8
60.94	156.1	148.4	93.47	598.4	576.8
64.41	182.7	176.2	95.88	654.7	628.8
72.64	261.8	252.6	100.25	767.6	737.0
77.65	322.9	311.0			

Nicol, 1886

mol%	Dp 70°	75°	80°	85°	90°	95°
0.99	3.9	5.2	6.8	7.8	9.4	11.1
1.96	7.5	9.7	12.2	15.0	18.2	21.2
2.91	10.0	12.6	15.7	19.7	24.0	29.3
3.85	13.4	16.7	20.8	26.0	31.7	39.0
4.76	15.8	20.2	25.0	30.8	37.9	46.5
9.09	27.3	34.6	43.0	53.7	66.2	81.3
13.04	36.5	46.2	57.9	71.8	88.6	108.8
16.67	45.0	56.8	71.4	88.4	109.4	134.2
20.00	51.2	65.7	82.5	102.0	125.3	153.9

Emden, 1887

t	p	t	p
4.78%		9.12%	
15.14	12.75	15.05	12.45
20.05	17.6	19.59	16.9
24.55	22.75	25.30	23.95
31.54	34.40	29.10	29.65
34.20	39.7	36.29	48.55
41.11	57.75	40.85	55.9
44.61	69.2	45.35	71.1
50.39	92.0	50.25	90.7
55.18	116.8	55.05	114.4
60.92	153.7	59.91	143.7
63.68	174.3	65.45	184.2
71.20	240.9	69.88	223.9
76.22	298.4	77.85	313.7
79.55	341.6	79.42	335.0
85.70	420.1	85.83	432.7
91.85	554.6	90.48	516.7
95.16	625.4	95.49	625.4
12.95%		16.72%	
19.11	15.7	12.20	10.0
25.67	23.35	20.87	17.65
28.77	29.65	25.86	23.7
35.74	41.65	29.65	29.6
39.99	52.6	34.88	39.8
45.44	69.35	39.87	51.8
51.00	92.3	45.18	68.2
56.38	119.85	49.70	85.9
61.14	149.6	54.74	110.1
66.80	192.9	59.56	137.3
70.27	224.45	65.81	183.7
75.57	283.65	70.25	222.15
80.25	392.4	73.86	259.1
85.42	421.75	81.27	351.9
90.05	503.4	85.60	417.1
94.18	587.9	89.33	481.75
		95.73	613.6
20.01%		20.01%	
17.71	14.3	59.80	137.9
25.43	22.9	64.47	170.0
31.34	31.8	69.94	216.5
35.10	39.3	76.20	282.8
40.35	52.65	79.52	323.15
45.93	70.1	85.97	419.0
48.50	80.4	91.05	509.5
55.20	111.2	95.36	597.4

Lincoln and Klein, 1907

%	p
25°	
0	23.76
9.46	23.09
17.88	22.49

Edgar and Swan, 1922

t	p	t	p
sat.sol.			
19	15.54	25	22.22
20	16.58	26	23.61
21	17.61	27	25.08
22	18.69	28	26.57
23	19.79	29	28.13
24	20.96	30	29.71

Frazer, Lovelace and Taylor, 1926

m	p	m	p
20°			
0.0979	17.4191	0.8205	16.4347
0.1947	17.3684	0.9114	17.0804
0.2913	17.3196	1.2029	17.0467
0.3875	17.2732	1.4652	16.9356
0.5415	17.2279	1.8098	16.8418
0.6261	17.2012	2.3504	16.7309
0.7631	17.1638	2.8873	16.5726
0.7631	17.1048		

Adams and Ulerz, 1929

t	p	t	p
sat.sol.			
10	8.87	30	28.84
15	12.26	40	48.67
20	16.21	50	78.56
25	21.94		

Legrand, 1835

%	b.t.	%	b.t.
10.87	101	65.02	110
20.88	102	67.66	111
29.68	103	69.97	112
37.34	104	72.04	113
43.91	105	73.91	114
49.54	106	75.62	115
54.34	107	77.02 sat.s.	115.9
58.43	108	77.07	116
61.96	109		

Kremers, 1856

Sat.sol. b.t. = 118°

De Heen, 1881

%	b.t.
8.32	101.2
11.44	101.5
18.26	101.8

Gerlach , 1886

%	b.t.	%	b.t.
0	100	50.24	106
6.98	100.5	54.65	107
13.19	101	58.59	108
18.70	101.5	61.12	109
23.67	102	65.33	110
28.06	102.5	68.25	111
32.20	103	70.83	112
35.90	103.5	73.26	113
39.21	104	75.37	114
42.20	104.5	77.22	115
45.05	105		

Baroni, 1893

%	b.t.	%	b.t.
1.74	100.160	10.21	100.937
3.70	100.342	12.24	101.136
6.00	100.545	14.39	101.356
7.93	100.724		

Buchanan, 1899

%	b.t.	%	b.t.
0	99.57		
3.55	99.92	11.52	100.59
5.04	100.00	12.55	100.70
6.06	100.10	14.68	100.91
7.23	100.19	16.72	101.11
8.33	100.29	21.48	101.61
9.47	100.39	25.48	102.13
10.62	100.49	42.66	102.84

Kahlenberg, 1901

%	b.t.	%	b.t.
2.71	100.25	25.26	102.63
3.59	100.38	26.22	102.76
5.71	100.53	29.35	103.15
7.78	100.74	30.80	103.33
8.87	100.84	33.07	103.65
11.55	101.12	34.80	103.87
12.89	101.23	36.39	104.12
16.49	101.63	38.14	104.39
17.58	101.74	38.58	104.45
21.32	102.17	41.44	104.76
21.91	102.25		

Berkeley and Applebey, 1911

sat.sol. b.t.= 115.549

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.0999	0.1023	1.211	1.539
.1454	.1504	.255	.612
.2031	.2126	.371	.819
.2158	.2264	.419	.893
.2467	.2617	.430	.917
.2585	.2747	.456	.954
.2661	.2831	.477	.993
.3350	.3620	.538	2.127
.3524	.3812	.543	.127
.4358	.4775	.591	.199
.4978	.5533	.629	.284
.5590	.6262	.687	.387
.6442	.7338	.759	.542
.6736	.7699	.868	.762
.7540	.8806	.990	3.011
.8072	.9513	2.066	.190
.8389	.9936	.084	.248
.8724	1.039	.136	.322
1.023	.251	.188	.469
.028	.267	.247	.578
.055	.298	.304	.740
.061	.318		
.109	.387		
.132	.423		

m_1= molality of potassium chloride

m_2= " " " nitrate

Robinson and Stokes, 1949

m	osmotic coefficient
	25°
0.1	0.906
.2	.873
.3	.851
.4	.833
.5	.817
.6	.802
.7	.790
.8	.778
.9	.767
1.0	.756
.2	.736
.4	.718
.6	.700
.8	.684
2.0	.669
2.5	.631
3.0	.602
3.5	.577

Rüdorff, 1861 and 1872

%	f.t.	%	f.t.
0.99	-0.25	7.41	-2.15
1.96	-0.55	9.09	-2.65
3.85	-1.1	10.71	-2.90
5.66	-1.6		

Gérardin, 1865

%	f.t.	%	f.t.
17.49	10	39.87	41
21.82	18	48.27	53
28.62	27		

Mulder, 1866

%	f.t.	%	f.t.
11.73	0	23.61	21
13.72	4.25	25.34	22.15
15.61	7	27.22	25
16.81	9.3	30.31	29.75
18.76	12.5	34.39	34.5
21.38	16.5	76.60	114.1 b.t.
22.12	18.5		

Nordenskjöld, 1869

%	f.t.
21.07	16.0
30.31	29.0
41.66	44.2

Guthrie, 1876 and 1878

%	f.t.	%	f.t.
0	0.0	40	41.0
1	-0.1	43.3	45.1
2	-0.3	49.2	54.7
3	-0.7	55.4	65.4
4	-1.1	62.5	79.7
5	-1.5	70.3	97.6
7	-2.2	71.56	100.0
8.5	-2.6	74.56	109.0
10	-2.9	74.97	114.0
11.2	-3.0 E	75.2	115.0
12	0.0	79.14	123.0
13	+2.0	84.67	151.0
15	6.0	89.94	201.0
20	14.0	95.11	262.0
25	21.0	98.86	300.0
35	35.0	100.00	320.0

Andreae, 1884

%	f.t.	%	f.t.
11.71	0	38.98	40
17.28	10	46.10	50
24.01	20	52.35	60
31.45	30		

Etard, 1894

%	f.t.	%	f.t.
79.8	139	86.0	190
83.7	158	89.0	215
83.9	160	90.4	225
84.0	175	91.6	258
84.2	180	96.5	283

Jaffé, 1903

%		f.t.
I	II	
28.59	23.78	20
32.48	27.17	25
36.75	30.79	30
44.63	39.02	40

Berkeley, 1904

%	f.t.	%	f.t.
11.84	0.40	52.64	60.05
20.49	14.90	61.02	76.00
32.25	30.80	67.76	91.65
42.69	44.75	75.71	114.0

Euler, 1904

27.43% f.t.=25°

Findlay, Morgan and Morris, 1914

%	f.t.
16.76	9.1
24.27	20.0
24.77	21.1
35.01	35.0

Tschugaeff and Chlopin, 1914

%	f.t.
46.39	50
51.55	58
53.41-53.87	62
57.04	68

Cornec and Hering, 1925

%	f.t.	%	f.t.
10.9	-2.89	46.1	50
12.0	0	60.5	75
27.7	+2.5	70.9	100

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
70.8	100	89.9	212
74.8	112	90.6	218
76.8	119	92.5	238
79.5	126	93.5	255
81.5	139	95.2	261
82.1	144	95.5	274
82.2	145	96.0	279
84.5	162	98.1	307
86.0	175	100	336
87.7	186		

Thompson and Vener, 1948

%	f.t.	%	f.t.
27.24	25	46.11	50
31.49	30	52.33	60
39.02	40	57.87	70

Ugaj and Simkina, 1950

%	f.t.	%	f.t.
0	0	14.0	+4.2
4.0	-1.1	17.7	10.0
8.0	-2.1	21.0	15.4
9.5	-2.5	25.0	21.2
10.2	-2.7 E	30.0	27.5
11.0	-0.8		

Akhumov and Pylkova, 1956

%	f.t.	t.spontan.crystal.
35.64	37	19
39.09	41	27
49.71	57	40.5
52.50	61	45
58.43	72	57
62.13	79	65.5
68.85	93.5	81
69.31	95.5	84
78.14	123	115
79.20	130	121.5
85.05	162	161

Jones, 1908

%	t.spontan.crystal.	%	t.spontan.crystal.
1.82	-1.5	33.22	28.9
3.80	-2.5	36.06	29.8
6.59	-3.2	37.54	33.9
9.27	-4.5	40.08	35.5
10.61	-5.3	41.07	38.4
12.33	-1.8	44.11	42.8
13.70	+0.7	44.89	42.6
16.62	4.4	48.93	47.6
17.98	8.1	50.21	51.8
19.84	10.0	54.15	55.6
23.27	16.0	54.87	60.5
23.32	12.9	57.55	60.9
24.85	16.4	58.69	66.1
28.47	22.2	61.33	71.9
28.67	20.1	63.90	76.6
28.74	23.5	64.41	75.9
32.71	27.0	65.98	82.8

Properties of phases

Bischof, 1850

%	d	%	d
18.75°			
2	1.0110	14	1.0901
4	.0238	16	.1041
6	.0367	18	.1184
8	.0497	20	.1328
10	.0630	22	.1477
12	.0765	24	.1629

Kremers, 1855 and 1861

%	d
19.5°	
0	0.9983
4.871	1.0290
9.618	.0599
14.044	.0901
17.965	.1179
21.488	.1436

t	d a	d b
19.5	1.0677	1.1256
40	.0591	.1154
60	.0485	.1037
80	.0361	.0905
100	.0219	.0760

Schiff, 1858, 1859 and 1860

%	d	%	d
21°			
0	0.9980	13	1.0797
1	1.0038	14	.0865
2	.0098	15	.0934
3	.0158	16	.1004
4	.0219	17	.1075
5	.0280	18	.1147
6	.0342	19	.1220
7	.0404	20	.1293
8	.0469	21	.1367
9	.0534	22	.1441
10	.0600	23	.1515
11	.0665	24	.1590
12	.0730	24.93	.1660 (satd)

Gerlach, 1859

%	d	%	d
15°			
0	0.99913		
1	1.00554	12	1.07811
2	.01195	13	.08501
3	.01835	14	.09191
4	.02477	15	.09881
5	.03117	16	.10604
6	.03780	17	.11329
7	.04443	18	.12052
8	.05105	19	.12777
9	.05769	20	.13500
10	.06431	21	.14261
11	.07121		

Schmidt, 1859

%	d	%	d
16°			
0	0.99897	12.51605	1.082177
4.25991	1.027142	16.79111	.112172
8.00925	.051653	20.36000	.138217

t	d	t	d
4.25991%		8.00925%	
14.44	1.027172	14.56	1.051702
17.86	.026961	18.10	.051305
21.65	.026760		
8.65025%		12.51605%	
14.44	1.055966	14.56	1.082250
24.16	.054871	18.10	.081666
26.29	.054626	22.00	.081112
16.79111%		20.36000%	
14.56	1.112261	16.01	1.137977
17.73	.111598	17.12	.137728
22.23	.110786	22.85	.136351
26.54	.109942	26.29	.135692
23.26709%			
20.91	1.158616		
24.16	.157816		
26.54	.157155		

Fouqué, 1867

%	d°	t	d
0.62	1.0046	22	1.0021
4.42	.0319	20	.0298
16.90	-	16	.1370

Buliginsky, 1868

%	d
	15°
0	0.99918
4.630	1.0270
7.572	.0465
10.405	.0657
13.218	.0862

Kohlrausch, 1879

%	d	%	d
	18°		
5	1.0305	20	1.133
10	1.0632	22	1.148
15	1.097		

Rönnberg, 1880

t	d	t	d	t	d
5%		10%		15%	
3.2	1.03047	3.2	1.05947	3.2	1.08722
9.7	.02955	10.0	.05837	10.0	.08563
15.2	.02852	15.0	.05718	15.0	.08411
19.5	.02751	19.5	.05590	19.5	.08262
24.5	.02593	24.5	.05422	24.5	.08078
20%		25%			
13.2	1.10982	14.4	1.13265		
18.5	.10802	21.5	.12992		
24.7	.10533	26.0	.12785		

Thomsen, 1881

mol %	d	mol %	d
	18°		
0	0.9986	2	1.0651
0.5	1.0173	4	1.1228
1	1.0336		

Volkmann, 1882

%	d	%	d
14°		15-16°	
7.31	1.0466	6.50	1.0411
13.66	.0900	12.03	.0784
18.68	.1263	19.94	.1347

Nicol, 1883

mol%	d 20°	d 40°
0.0	0.99823	0.99224
0.5	1.01550	-
1.0	.03190	1.02432
2.0	.06335	.05505
3.9	.12065	-
4.8	.14684	1.13667

Nicol, 1884

t	d
sat.sol.	
75	1.17103
85	.15315
95	.07505

Kanonnikov, 1885

%	t	d
0	21.0	0.9980
14.88	21.4	1.0960
16.81	20.9	1.1094

Le Blanc, 1889

%	d
	20°
0	0.99823
24.53	1.13099

Bodländer, 1891

c	%	d
	17.7°	
24.85	21.63	1.1475
25.12	21.89	.1465
	20.5°	
27.68	23.81	1.1625

Barnes and Scott, 1898

%	d	%	d
	20.1°		
0	0.9982	8.706	1.0553
0.741	1.0030	13.93	.0913
2.030	.0113	17.88	.1200
2.848	.0165	21.95	.1510
4.389	.0264	25.54	.1783
5.393	.0331		

Pann, 1901

%	d 10°	d 18°	d 30°
0	0.999727	0.998622	0.995673
5.92	1.03845	1.03646	-
12.31	.08170	.07876	1.06994
16.39	.10778	.10716	-
20.80	-	.13884	1.13771
28.12	-	-	.19255

Berkeley, 1904

%	t	d
	sat.sol.	
11.84	0.40	1.0817
20.49	14.90	.1389
32.25	30.80	.2218
42.69	44.75	.3043
52.64	60.05	.3903
61.02	76.00	.4700
67.76	91.65	.5394
75.71	114.0	,6269

Euler, 1904

sat.sol. (25°) 27.43% d= 1.190

Loewenfeld, 1905

%	d
15°	
0	0.9993
5	1.0313
9.3	.0594
13.7	.0883
18.1	.1132

Zecchini, 1905

%	t	d
12.157	10.4	1.08038
14.656	14	.09638
21.805	16.4	.14828

Chèneveau, 1907

%	d	%	d
	15°		
0	0.9991	12.51	1.0816
2.23	1.0140	14.42	.0953
4.40	.0276	16.25	.1089
6.50	.0417	18.06	.1219
8.54	.0552	19.84	.1336
10.55	.0695		

21.08% 16.3° d= 1.1420

Getman , 1907

c	%	d
	18°	
5	4.845	1.0306
10	9.43	.0596
15	13.78	.0886
20	17.89	.1176
25	21.73	.1504

Getman and Wilson, 1908

%	d	%	d
	21°		
0	0.9980	5	1.0300
1	1.0065	10	.0621
2	.0131	15	.0956
3	.0197	20	.1316
4	.0264		

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	0.4638	1.02814
0.0937	1.00480	0.925	.05712
0.1870	.01211	2.3496	.13898

Chernai, 1912

%	t	d
11.71	0	1:084
17.27	10	1.120
23.44	20	1.161
31.44	30	1.212
38.88	40	1.282
46.10	50	1.339
52.37	60	1.403
57.97	70	1.446
62.82	80	1.490
67.10	90	1.510

Stocker, 1920

%	t	d
0	20.7	0.9981
5.75	18.2	1.0348"
15.19	18.7	1.0985
21.46	20.3	1.1428

Manchot, Jarhrstorfer and Zepter, 1924

c	d
25°	
10.313	1.0586
10.8794	1.0618
21.739	1.1231
21.7589	1.1232

Cornec and Hering, 1925

%	t	d
27.7	25	1.189
46.1	50	1.332
60.5	75	1.467
70.9	100	1.569

Hrynakowski, 1927

sat.sol. (50.1°) d= 1.3375

Geffcken, 1929

m	d	m	d
25°			
0	0.99703	1.3913	1.07650
1.2520	1.07287	3.0458	.15677
1.3544	1.07452	3.0470	.15686

Spacu and Popper, 1934

%	d
20°	
0	0.99823
10.498	1.06553
15.163	.09748
19.760	.13043
21.016	.13931

Gibson, 1935

%	d
25°	
0	0.9971
8.16	1.0488
13.68	.0858
20.22	.1319
26.53	.1787

Guillaume, 1946

%	t	d
9.9	20	1.0638
14.2	20.5	.0935
18.0	21.5	.1208
22.0	20	.1494
22.5	20	.1527

De Heen, 1881

t	r.v.*	t	r.v.	t	r.v.
10.00	1.000000	10.00	1.000000	10.00	1.000000
16.35	.002273	12.95	.000833	15.13	.001339
20.17	.003637	18.42	.002489	20.42	.002907
29.26	.007411	25.41	.004900	28.52	.005653
36.70	.010775	32.47	.007766	35.59	.008435
48.71	.016886	40.57	.011369	42.29	.011388
56.06	.021048	49.24	.015696	50.39	.015365
61.80	.024497	57.97	.020683	57.45	.019133
70.38	.029650	64.16	.024082	62.51	.022025
		70.00	.027881	70.87	.027185

* r.v.= relative volume

Nicol, 1884, 1886 and 1887

mol%	molar volume
20°	
1	1839.18
2.9	1920.83
4.7	2007.74

t	r.v.	t	r.v.
20°			
1 mol%		2.9 mol%	
20	100.000	20	100.000
45.2	100.933	45.4	101.201
50.4	101.182	50.2	101.453
56.2	101.473	56.5	101.801
61.6	101.763	61.3	102.078
67.8	102.107	67.2	102.433
72.2	102.368	72.4	102.748
78.1	102.736		
4.7 mol%			
20	100.000	61.3	102.078
45.4	101.201	67.2	102.433
50.2	101.453	72.4	102.748
56.5	101.801		

Schmidt, 1859

%	t	π
21.8	19.3	35.4
17.0	18.9	37.5
11.7	18.7	40.1

Gibson, 1935

%	π (1-1000 bars)
0.00	39.35
8.16	36.63
13.68	34.82
20.22	32.83
26.53	30.92

Pohl, 1852

%	relative π
13.9°	
0	1.000
9.30	0.900
15.64	.827

Getman, 1907

c	η
18°	
5	1046
10	1044
15	1043
20	1053
25	1076

c	η 10°	η 30°	η 50°
6.54	1260	790	570
13.68	1230	790	580
20.70	1230	810	590

Hrynakowski, 1927

sat.sol. (50.1°) η=900

Tammann and Rabe, 1928

%	η 0°	η 10°	η 30°	η 45°
9.17	1623.6	1216.7	792.5	401.5
16.81	-	1212.1	811.4	424.2
23.28	-	-	846.2	449.2
28.83	-	-	885.1	474.6
33.59	-	-	-	502.2

Buliginsky, 1868

%	σ
15°	
0	73.26
4.630	73.79
7.572	74.19
10.405	74.48
13.218	74.84

Volkmann, 1882

%	σ	%	σ
14°		15-16°	
7.31	74.7	6.50	74.5
13.66	75.2	12.03	74.8
18.68	76.0	19.94	75.9

Pann, 1901

%	σ 10°	σ 30°
0	74.02	71.02
5.92	74.75	-
12.31	75.37	72.21
16.39	75.95	-
20.80	-	73.53
28.12	-	74.62

Forch, 1905

N	t	σ
0.875	15.7	77.69
1.260	15.6	78.22
1.752	15.9	78.61
2.175	15.7	79.15

Loewenfeld, 1905

%	σ
	15°
0	74.21
5	73.61
9.3	74.74
13.7	75.56
18.1	74.76

Kleine, 1908

N	σ (water=100)	N	σ (water=100)
	18°		
0	100	1.005	101.550
0.141	100.258	.170	101.756
.233	100.418	.282	101.961
.441	100.734	.429	102.120
.747	101.227	.556	102.312
.763	101.239	.574	102.199
.770	101.234		

Stocker, 1920

%	σ
	18°
0	72.59
5.75	73.07
15.19	74.46
21.46	75.41

Hrynakowski, 1927

Sat.sol. 50° σ=74.80°

Börner, 1869

t	$n_{H\alpha}$	t	$n_{H\beta}$	t	$n_{H\gamma}$
		10%			
45.2	1.336206	44.7	1.342511	43.95	1.345888
40.9	.336946	40.4	.343194	40.3	.346677
35.85	.337722	35.9	.343913	35.4	.347395
30.1	.338569	30.05	.344919	30.0	.348345
		20%			
44.7	1.342979	44.6	1.349455	44.1	1.353228
39.95	.343734	40.3	.350332	40.55	.353818
35.25	.344632	35.85	.351084	34.65	.354871
29.5	.345601	30.15	.352085	29.3	.355834
		30%			
43.7	1.349187	44.7	1.355798	44.3	1.359645
38.35	.350189	39.1	.356904	39.9	.360464
33.75	.351119	35.3	.357652	35.1	.361459
30.6	.351638	30.7	.358471	30.5	.362383

Kanonnikov, 1885

%	n Hα	n D	n Hβ
		21°	
0	1.33120	1.33300	1.33728
14.88	.345060	.347100	.351906
16.81	.347094	.349125	.353968

Le Blanc, 1889

%	n_D
	20°
0	1.33325
24.53	.35213

Borgesius, 1895

%	t	$D(n_D)$
5.32	20.9	0.000606
20.59	21.7	0.002434

Zecchini, 1905

%	t	n_D
12.157	10.4	1.34612
14.656	14.0	.34811
21.805	16.4	.35449

Chéneveau, 1907

%	n_D	%	n_D
	15°		
0	1.3334	12.51	1.3456
2.23	.3357	14.42	.3473
4.40	.3377	16.25	.3491
6.50	.3397	18.06	.3510
8.54	.3417	19.84	.3527
10.55	.3437		

t	%	n C	n D	n Tl	n F	n G'
16.3	21.08	1.35160	1.35364	1.35560	1.35837	1.36212

Getman and Wilson, 1908

%	n_D	%	n_D
	21°		
0	1.33288	5	1.33773
1	.33426	10	.34240
2	.33521	15	.34735
3	.33600	20	.35263
4	.33690		

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.4638	1.33768
0.0937	.33419	0.925	.34179
0.1870	.33507	2.3496	.35324

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	0.5	1.33801
0.1	.33426	1.0	.34244
0.2	.33520	2.0	.35066

Geffcken, 1929

m	n_{He}	m	n_{He}
	25°		
0	1.332590	1.3913	1.344026
1.2520	.343548	3.0458	.354964
1.3544	.343777	3.0470	.354976

Spacu and Popper, 1934

%	n_{He}
20°	
0	1.3324865
10.498	.342317
15.163	.346781
19.760	.351313
21.016	.352443

Guillaume, 1946

%	t	n_{5780}
9.9	20	1.0638
14.2	20.5	.0935
18.0	21.5	.1208
22.0	20	.1494
22.5	20	.1527

%	t	* $\alpha_{magn.}10^6$ (5780)
9.9	20	3.707
14.2	20.5	.586
18.0	21.5	.480
22.0	20	.370
22.5	20	.354

* in radians, gauss, centim.

Okazaki, 1933

%	Verdet's constant (3441 Å)
28°	
5.66	0.04385
12.15	.04366
19.83	.04275
22.93	.04225

Kohlrausch, 1879

%	κ
	18°
5	452
10	833
15	1181
20	1499
22	1629

Kramers, 1898

%	κ 50°	100°	150°	200°	250°	300°	350°
5.09	790	1320	-	-	-	-	-
9.33	1430	2330	-	-	-	-	-
12.89	1930	3120	-	-	-	-	-
15.95	2340	3680	-	-	-	-	-
18.58	2670	4200	-	-	-	-	-
20.88	2960	4680	-	-	-	-	-
22.91	3230	5110	-	-	-	-	-
24.71	3450	5440	-	-	-	-	-
26.32	3640	5690	-	-	-	-	-
27.30	3770	5880	-	-	-	-	-
30.25	3970	6150	-	-	-	-	-
32.32	4070	6310	-	-	-	-	-
34.08	-	6360	-	-	-	-	-
35.57	-	6360	-	-	-	-	-
36.87	-	6280	-	-	-	-	-
39.49	-	5950	7880	-	-	-	-
41.40	-	5130	7200	-	-	-	-
44.10	-	-	6010	6480	-	-	-
45.92	-	-	4080	5200	-	-	-
47.20	-	-	-	-	5690	-	-
48.14	-	-	-	-	-	7090	-
48.91	-	-	-	-	-	6310	7180
49.51	-	-	-	-	-	5670	6540

Dennhardt, 1899

N	λ 0°	10°	18°
0.5	560	716	886
1	509	662	798
2	-	576	679
2.5	-	-	635

Lorenz, 1912

N	t	λ
18.27	333	33.16
18.18	343	35.50
18.12	353	27.38
18.06	363	39.98
18.00	373	40.66
17.94	383	42.36
17.88	393	44.07
17.82	403	46.01
17.72	418	47.62

Thermal constants

Schüller, 1869

%	U
	at room temperature
9.09	0.9182
16.67	0.8589
23.08	0.8090

Winkelmann, 1873

%	U	%	U
3.05	0.9673	11.11	0.8997
4.15	.9575	15.31	.8721
5.62	.9458	19.80	.8484
8.40	.9206		

Thomsen, 1881

mol%	U
	18°
0.5	0.966
1	.942
2	.901
4	.832

Richards and Gucker, 1925

mol%	t	U
3.85	18.01	0.83284

Randall and Rossini, 1929

m	U	m	U
	25°		
0.00	0.9979	0.20	0.9763
.01	.9967	.35	.9615
.02	.9956	.50	.9478
.05	.9922	.75	.9267
.10	.9867	1.00	.9077

Rutskov, 1948

%	U 25°	U 50°	U 75°
4.812	-	0.9463	0.9520
9.183	-	.9090	.9122
21.91	0.8015	.8020	.8030
41.23	-	.6575	.6565

Winkelmann, 1873

%	Q diss(cal/g nitrate) 0°	50°
3.05	-89.43	-72.4
4.15	87.16	71.9
5.62	83.97	71.1
8.40	83.0	69.8
11.10	81.3	69.2
15.30	76.9	66.8
19.80	73.4	65.5

Staub, 1890

%	Q diss (cal/g nitrate)
2.347	-90.99
6.103	-83.40
9.080	-81.13
11.432	-78.06
13.116	-76.21

Scholz, 1892

N	Q diss (cal/g nitrate)
0.0625	-93.55
.125	-92.99
.25	-90.75
.5	-87.55
1.0	-82.00
1.3	-79.42

Varali-Thevenet, 1902

%	Q diss (cal/g salt)	2^{nd} series
0.2	-95.1	-
0.4	91.6	-
0.6	90.0	-
0.8	88.7	-
1	87.5	-
2	82.9	-90.5
4	80.2	82.3
6	74.5	76.2
8	68.2	75.0
10	66.8	74.4
12	63.5	53.0
14	26.1	4.6 ?

Mondain-Monval, 1923

Q diss. 0.8-0.9 mol% 0° -8800 cal/mole nitr.
" 18° -8300

Q dol sat.sol. 0.8 mol % 0° -1470
18° -1710

Q mix (large amount of sat.sol. + 1 mole aqua
0° -20
18° -50

Q diss (large amount of sat.sol. + 1 mole salt)
0° -870
18° -1140

Q mix limit
0° -6400
18° -5600

Bishop, 1850

m initial	m final	Q dil(by mole nitrate) a	b*
2.380	2.049	-129.9	-129.9
2.028	1.750	119.1	264.1
2.380	.550	341.1	341.1
1.750	.500	116.9	381.0
.500	.280	110.2	491.2
.576	0.637	397.5	743
.578	.493	462.9	808
2.380	.291	920	920

* initial molality= 2.380

Hunter and Bliss, 1944

mol%	Q vap (cal/g water)	Q dil (cal/mole water)
	30°	
0.0198	-578.4	-32.2
.0200	579.3	-16.0
.0436	577.8	-43.9
.0439	577.0	-58.8
.0310	578.3	-39.2
.0315	577.9	-42.9

Water (H_2O) + Potassium perchlorate ($KClO_4$)

Carlson, 1910

%	f.t.	%	f.t.
0.78	0	8.01	60
1.76	20	12.87	80
4.59	40	17.34	100

Calzolari, 1912

%	f.t.
1.86	20.5
10.95	70
18.16	99

Cornec and Neumeister, 1929

%	f.t.	%	f.t.
0.75	0	10.36	75
2.03	25	18.17	100
4.91	50		

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
23.1	117	50.0	192
26.1	125	52.5	198
26.9	128	56.3	210
33.1	142	60.7	227
35.2	146	63.4	236
43.1	174	65.6	244
46.0	180	69.9	265
46.8	180	100	610

Carlson, 1910

%	t	d	%	t	d
0.78	0	1.007	8.01	60	1.033
1.76	20	.011	12.87	80	.053
4.59	40	.022	17.34	100	.067

Cornec and Neumeister, 1929

%	t	d
0.75	0	1.005
2.03	25	.013
4.91	50	.017
10.36	75	.036
18.17	100	.068

Hantzch and Düringen, 1929

%	d	%	d
	20°		
2.2130	1.007805	25.332	1.16227
3.9910	.02576	35.9449	.25246
7.6938	.040516	52.927	.43656
13.174	.07520	58.486	.50704
15.7425	.09261	69.570	.65994

%	n_D	%	n_D
	20°		
2.2130	1.33435	25.332	1.35240
3.9910	.33648	35.9449	.363426
7.6938	.338124	52.927	.38698
13.174	.34211	58.486	.39641
15.7425	.34420	69.370	.41542

Water (H_2O) + Potassium perrhenate ($KReO_4$)

Hölemann and Kleese, 1938

%	f.t.	%	f.t.
0.61	10.5	50.7	220
1.185	25	59.9	239
3.19	50	71.9	290
4.35	60	80.1	335
6.95	75	84.6	366
8.985	85	89.3	401
12.16	99	94.4	445
12.6	102	95.4	456
14.0	112	96.8	470
26.3	154	97.4	498
39.7	194	100.0	518

Water (H_2O) + Potassium permanganate ($KMnO_4$)

Voerman, 1905

%	f.t.	%	f.t.
0.58	-0.18	4.01	+10
0.99	-0.27	4.95	15
1.98	-0.48	7.00	25
2.91	-0.58 E	10.40	40
		14.35	50

Baxter, Boylston and Hubbard, 1906

%	f.t.	%	f.t.
2.75	0	11.16	40.0
4.13	9.8	12.73	45.0
5.96	19.8	14.45	50.0
7.06	24.8	16.20	55.0
8.28	29.8	20.02	65.0
9.64	34.8		

Worden, 1907

%	f.t.	%	f.t.
3.15	4	11.14	40
4.21	10	12.70	45
5.12	15	14.35	50
5.22	15.56	16.15	55
6.11	20	18.03	60
7.11	25	19.99	65
8.31	30	22.24	70
9.62	35	24.44	75

Water (H_2O) + Potassium orthophosphate(K_3PO_4)

Jänecke, 1927

%	f.t.	
44.2	7.5	
49.0	23.3	
57.0	43.2	
59.7	45.1	(8+1)

Ravich, 1938

wt%	mol%	f.t.	
4.54	0.40	-1.18	A
9.75	0.91	-2.60	
15.43	1.52	-4.6	
21.74	2.30	-7.7	
27.34	3.09	-12.0	
31.53	3.76	-15.8	
35.12	4.39	-20.0	
38.33	5.0	-24.0	E A + (9+1)
40.25	5.40	-28.2	E A + (7+1)
42.92	6.0	-8.8	(7+1)
44.26	6.31	0	(7+1)
46.83	6.95	10.0	"
49.62	7.71	20.0	"
51.42	8.23	25.0	"
53.08	8.75	30.0	"
55.43	9.54	35.0	"
57.51	10.30	40.0	"
59.46	11.06	42.6	"
60.84	11.64	44.5	"
61.94	12.13	45.4	"
62.51	12.39	45.6	"
63.12	12.68	45.6	"
-	-	45.4	(7+1) + (3+1)
63.17	12.70	25.0	
63.19	12.71	30.0	
63.33	12.77	35.0	
63.41	12.81	40.0	(3+1)
63.56	12.89	45.0	
63.80	13.0	50.0	
64.08	13.14	60.0	
47.62	6.21	0.0	
49.80	7.16	5.0	(9+1)
52.23	8.49	8.8	
57.72	10.0	12.3	

Moore, 1895-96

N	d	η (water=1)
	18°	
0.00	0.9987	1.000
0.125	1.0227	.070
0.25	.0471	.126
0.50	.0933	.298
1.0	.1805	.759

Water (H_2O) + di-Potassium orthophosphate (K_2HPO_4)

Ravich, 1938

wt%	mol%	f.t.	
16.78	2.04	-4.2	A
23.60	3.09	-6.4	"
29.61	4.17	-9.0	"
34.10	5.07	-11.7	"
36.78	5.67	-13.5	E A + (6+1)
46.11	8.12	0	(6+1) J
50.12	9.40	+4.95	"
54.43	10.90	9.7	"
57.89	12.44	13.15	"
-	-	14.3	(6+1) + (3+1)
60.82	13.82	14.6	(6+1) II
61.73	14.29	14.85	"
62.96	14.94	14.7	"
65.95	16.68	12.8	"
60.09	18.75	8.2	"
57.05	12.01	0	(3+1) II
59.08	12.96	10.0	"
60.16	13.49	15.0	" I
61.52	14.16	20.0	"
62.74	14.83	25.0	"
64.13	15.60	30.0	"
65.68	16.51	35.0	"
67.54	17.68	39.5	"
69.83	19.29	44.0	"
71.26	20.42	46.0	"
72.64	21.55	51.0	B
72.50	21.38	56.0	"
72.79	21.66	63.0	"

Selva, 1947

%	f.t. (6+1)	f.t. (3+1)
16.44	-4.2 A	-
25.92	-7.4	-
34.05	-11.6	-
36.67	-13.9 E	-
38.80	-10	-
45.58	0	-
53.94	+10.2	-
57.28	-	0
57.63	13.4	-
58.86	-	10.2
59.76	-	13.4
60.20	14.7	14.7 tr.t.
60.30	-	17.2
62.03	-	25
64.56	-	34
68.95	-	44.2
70.76	-	46.2
71.50	44.2	-
71.63	28.7	-
71.90	-	48.3 E
72.04	51	-
72.53	64.8	-
72.83	84.5	-
73.80	99.4	-

Forch, 1905

t	vol rel.	t	vol rel.
	0.5 N		
0.02	1.000000	24.99	1.005836
4.98	1.000759	30.02	1.007564
10.03	1.001750	34.96	1.009412
15.00	1.002924	40.01	1.011454
20.02	1.004290		
	1 N		
0.02	1.000000	19.94	1.005790
4.95	1.001243	25.00	1.007589
10.01	1.002645	30.03	1.009485
14.99	1.004154	35.02	1.011467
34.93	1.000000	39.98	1.002125
	2 N		
0.00	1.000000	19.97	1.007133
5.01	1.001695	24.97	1.009082
10.01	1.003438	30.02	1.011127
14.96	1.005236		
29.98	1.000000	40.03	1.004221
34.95	1.002053		

Moore, 1895-96

N	d	η (water=1)
	18°	
0.00	0.9987	1.000
0.125	1.0167	.039
0.250	.0343	.095
0.50	.0700	.206
1.00	.1383	.542
2.00	.2633	2.309

Hrynakowski, 1927

t=50.70° d=1.2216 η =1299 σ =69.71

Water (H_2O) + mono-Potassium orthophosphate (PO_4H_2K)

Tammann, 1885

%	p	%	p
	100°		
11.91	740.8	40.37	687.2
19.16	729.8	47.21	667.7
32.36	705.2	47.30	667.3

Adams and Merz, 1929

t	p	t	p
	sat.sol.		
10	8.96	30	29.60
15	12.62	40	51.46
20	16.89	50	85.63
25	22.76		

Jones, 1904; Jones and Getman, 1904

N	f.t.	N	f.t.
0.50	- 1.525	1.00	- 2.780

Apfel, 1911

N	f.t.	N	f.t.
0.77	0	50	2.15
1.465	7	70	2.695
.47	25	83	3.04
.46	25		

Kazantsev, 1938

%	f.t.	%	f.t.
12.09	-2.75	21.77	30
12.79	0	23.51	35
14.05	+5	25.31	40
14.63	7	27.17	45
15.46	10	29.07	50
16.93	15	33.01	60
17.84	18	37.10	70
18.46	20	41.29	80
19.44	23	42.64	83
20.09	23	45.63	90

Bergman and Bochkarev, 1938

%	f.t.	%	f.t.
4	+0.7	16	+13.6
8	-1.5	18	19.8
10	-2.1	20	26.0
12	+0.8	22	31.8

E: 11.7% -2.6°

Polosin and Ozolin, 1940

%	f.t.	%	f.t.
0	0	13	2.63
2	-0.34	13	3.8
4	-0.83	14	6.95
6	-1.2	15	10.08
8	-1.6	16	13.26
10.5	-2.03	18	19.6
11	-2.23	20	25.97
12	-2.48	22	32.08
12	-2.57	24	38.3
12.5	+1.1		

Ugaj, 1950

%	f.t.	%	f.t.
3	-0.5	15	10.2
6	-1.1	18	20.0
9	-1.8	20	26.0
11	-2.3	22	32.0
11.6	-2.5 E	25	39.5
13	+3.4		

Robinson and Stokes, 1949

m	osmotic coefficient
25°	
0.1	0.901
.2	.868
.3	.843
.4	.823
.5	.805
.6	.789
.7	.773
.8	.760
.9	.747
1.0	.736
.2	.716
.4	.698
.6	.683
.8	.669

Kohlrausch, 1879

%	d
18°	
5	1.0341
10	1.0691
15	1.1092

Moore, 1895 - 1896

N	d	N	d
18°			
0.00	0.9987	0.5	1.0442
0.25	1.0220	1.0	1.0885"

Chomojakow, Jaworowskaja and Schirokich, 1933

%	d
23°	
19.05	1.2168
19.41	1.2288
sat.sol.	1.2331

Mason and Culvern, 1949

N	d	N	d
25°			
0.004	0.99768	0.07	1.0038
0.005	0.99776	0.08	1.0049
0.006	0.99786	0.09	1.0057
0.007	0.99796	0.1	1.0062
0.009	0.99806	0.2	1.0161
0.01	0.99776 (?)	0.3	1.0252
0.02	0.99872	0.4	1.0344
0.03	1.0001	0.6	1.0523
0.04	1.0010	0.8	1.0696
0.05	1.0021	0.9	1.0777
0.06	1.0030	1.0	1.0870

Moore, 1895 - 1896

N	η	N	η
18°			
0.00	1.000	0.5	1.146
0.25	1.057	1.0	1.306

Jones, 1904 and Jones and Getman, 1904

N	λ	N	λ
0°			
0.05	51.00	0.50	38.24
0.10	46.71	1.00	32.21
0.20	42.45	1.50	28.98

Mason and Culvern, 1949

N	λ	N	λ
25°			
0.001	103.715	0.07	87.993
0.001401	104.015	0.08	86.104
0.002	102.177	0.09	86.015
0.002564	100.694	0.1	85.641
0.004	99.896	0.2	79.279
0.005	99.278	0.3	75.119
0.006	98.525	0.4	71.826
0.007	98.016	0.5338	68.200
0.008	97.720	0.6	66.627
0.009	97.237	0.7	64.017
0.01	96.936	0.8	62.059
0.02	94.149	0.9	60.159
0.03	92.066	1.0	58.581
0.04	90.578	1.1445	56.415
0.05	90.209	1.885	48.391
0.06	88.046		

Jones, 1904 and Jones and Getman, 1904

N	n_D	N	n_D
0°			
0	1.33395	0.50	1.33272
0.05	.32593	1.00	.33978
0.10	.32711	1.50	.34623
0.20	.32829		

Gladstone and Hibbert, 1897

%	(D) molecular refraction
10.6	29.7
19.7	29.47

Chomojakow, Jaworowskaja and Schirokich, 1933

mol%	Q mix	mol%	Q mix
0.35	-4640	1.72	-4416
0.70	-4574	2.06	-4364
1.04	-4515	2.40	-4320
1.39	-4464	2.74	-4277

Water (H_2O) + Mono-Potassium orthophosphate. phosphoric acid ($KH_5P_2O_8$)

Parravano and Mieli, 1908

%	f.t.	%	f.t.
0	0	68.44	65.2
3.337	-0.6	72.43	78.0
8.284	-1.7	77.60	87.5
12.13	-2.5	85.88	105.5
20.50	-5.7	92.18	120.0
29.43	-6.7	95.73	134.5
36.98	-9.2	96.10	135.0
45.80	0.	98.85	137.5
50.33	+10.9	100.00	139.0

Water (H_2O) + Potassium pyrophosphate ($K_4P_2O_7$)

Morey and Chen, 1956

t	P
V+L+C	
374	20
400	59
500	148
600	165
700	227

Water + Potassium orthoarsenate (K_3AsO_4)

Dupuy, 1884

%	f.t.
21.94	15
24.37	30
30.72	60
44.25	100

Water + Potassium acid arsenate (KH_2AsO_4)

Tammann, 1885

%	p	%	p
100°			
12.66	743.0	39.85	695.3
19.71	733.3	41.19	692.3
29.34	717.4	45.43	680.3

Water + Potassium thioantimonate (K_3SbS_4)

Dank, 1908

%	f.t.	%		f.t.
9.5	-1.3	62.0 E		-34
17.1	-2.6	65.5	(6+1)	-10
24.2	-4.0	69.1		-4.5
30.7	-5.6	75.4		0
35.4	-7.2	76.2	(5+1)	+10
42.9	-10.6	75.1		30
48.8	-13.5	77.7	(3+1)	50
52.6	-18.5	79.2		80
59.6	-28.8			

Water (H_2O) + Potassium tetraborate ($K_2B_4O_7$)

Keshan, 1955

%	f.t.	%	f.t.
8.2	0	18.0	25.0
13.0	14.5	22.8	34.7
15.5	19.6	35.3	56.8

Water + Potassium pentaborate (KB_5O_8)

Rollet and Andres, 1931

%	f.t.	%	f.t.
1.54 E	-0.53	13.22	75
1.56	0	13.62	76.65
1.77	5	15.5	82.3
2.66	18	17.02	87.15
3.8	30	18.0	89.8
5.72	45	19.85	94.8
8.45	57.6	22.3	100
9.85	62.8	23.0	101.65
11.5	69	23.4	102.2 b.t.
22	100 (8+1)	70	230
45	150	80	300 (2+1)
63	170 tr.t.		

Water (H_2O) + Potassium acid iodate (KHI_2O_6)

Rosenheim and Liebknecht, 1899

%	b.t.	%	b.t.
0	100.000	5.72	100.260
2.52	100.143	7.63	100.360
3.30	100.157	12.85	100.605
5.51	100.264	18.93	101.025

Water (H_2O) + Potassium carbonate (K_2CO_3)

Heterogeneous equilibria

Nicol, 1884

t	p
	sat.sol.
75	123.5
85	172.7
95	244.8

Tammann, 1885

%	p	t	p
	100°		
11.98	730.3	38.08	585.8
18.33	707.8	39.91	561.9
25.63	674.7	46.38	491.7
35.66	606.9	53.34	398.4
36.45	600.1	59.34	341.9

Tammann, 1885

t	p 0%	18.91%	30.77%	42.38%	51.82%
20.51	18.1	16.8	15.1	11.8	8.6
27.46	27.5	25.7	22.9	17.7	12.9
31.81	35.3	32.8	29.3	23.0	16.6
38.00	49.7	45.7	41.2	33.0	23.4
40.85	57.9	52.6	47.6	38.3	27.6
45.00	71.9	65.9	59.2	37.9	34.7
49.70	91.2	83.8	75.9	61.0	44.3
51.97	102.0	93.6	84.8	68.6	49.9
54.48	115.2	105.8	95.2	77.6	56.6
57.05	130.2	119.9	108.2	88.3	64.3
59.35	145.0	133.8	120.8	98.4	72.7
62.15	165.0	153.0	138.2	112.5	82.9
65.04	188.0	174.6	157.6	128.5	95.4
68.37	217.8	201.8	182.9	149.2	111.2
70.03	234.1	217.2	196.2	161.0	120.4
72.86	264.3	244.5	221.3	181.8	136.0
74.25	280.3	259.7	235.0	193.6	145.0
76.30	305.4	282.9	255.5	210.8	158.2
77.89	326.2	302.5	273.9	226.0	169.9
80.58	363.9	338.0	306.4	252.9	190.9
82.86	398.6	370.4	336.2	278.8	211.2
84.76	429.8	399.7	363.8	301.7	229.6
87.49	477.9	444.7	403.6	335.4	256.5
89.29	511.8	476.5	432.8	359.5	-
91.65	559.8	521.0	473.2	393.2	-
94.77	628.7	585.2	532.2	443.8	-
99.86	756.2	-	638.1	536.1	-

Lescoeur, 1896

t	p (2+1)	sat.sol.
20	-	7.3
25	-	10.0
30	-	13.1
35	-	16.25
40	-	20
50	-	44
60	-	65
80	-	147
90	57	-
100	100	300
110	172	-
120	275	-

Legrand, 1835

%	b.t.	%	b.t.	
11.50	101	50.67	115	
18.37	102	51.80	116	
23.67	103	52.89	117	
27.96	104	53.93	118	
31.56	105	54.95	119	
34.68	106	55.94	120	
37.34	107	57.81	122	
39.72	108	59.52	124	
41.83	109	61.17	126	
43.69	110	62.64	128	
45.35	111	64.04	130	
46.87	112	65.37	132	
48.24	113	66.63	134	
49.49	114	67.28	135	sat.sol.

Kremers, 1856

sat.sol. b.t.= 135°

Gerlach, 1886

%	b.t.	%	b.t.
0	100	54.01	118
9.31	101	55.05	119
18.37	102	56.04	120
24.24	103	56.97	121
28.57	104	57.90	122
31.20	105	58.76	123
35.27	106	59.58	124
37.89	107	60.39	125
40.12	108	61.24	126
42.20	109	62.05	127
43.98	110	62.89	128
45.50	111	63.71	129
46.95	112	64.37	130
48.33	113	65.21	131
49.62	114	65.92	132
50.85	115	66.61	133
50.85	116	66.95	133.5
53.16	117		

Jones, 1904; Jones and Getman, 1904

N	b.t.	N	b.t.
0.05	100.06	0.4	100.38
0.1	100.10	1.0	101.08
0.2	100.19	2.0	102.67

Poggiale, 1843

%	f.t.	%	f.t.
45.39	0	54.39	60
47.02	10	55.96	70
48.47	20	57.31	80
50.02	30	58.87	90
51.50	40	60.59	100
53.02	50	67.22	135

Rüdorff, 1861

%	f.t.	%	f.t.
1.39	-0.45	7.17	-2.45
2.97	-0.95	10.87	-3.9
4.79	-1.7	12.93	-4.7

Mulder, 1866

%	f.t.	%	f.t.
47.20	0	53.00	26
50.58	4.25	54.51	47.25
51.80	8.5	55.81	57.5
52.31	13.5	58.50	81
52.33	14	60.80	99

de Coppet, 1872

%	f.t.	%	f.t.
1.96	-0.6	16.67	-6.75
5.66	-1.9	19.09	-8.25
6.98	-2.4	20.00	-8.85
9.09	-3.2	23.08	-11.15
10.71	-3.85	25.93	-13.8
12.28	-4.55	28.57	-16.7
13.94	-5.35	33.33	-23.45
15.25	-6.0	37.50	-31.6

Jones, 1904; Jones and Getmann, 1904

N	f.t.
0.40	-1.683
1.00	-4.375
2.00	-9.710

Meyerhoffer, 1905

%	f.t.
39.6 E	-36.5
51.2 (2+1)	0
53.2 "	+25

Rubtsov, 1918

%	f.t.
52.01	16.5
52.80	19.5

Mc Bain, 1927

53.9% (sat.sol.) f.t.= 40° (3+2)

A.E Hill and D.G. Hill, 1927

%	f.t.
52.74	25 (3+2)
53.66	36 "

Hill and Miller, 1927

%	f.t.
52.8	20
53.2	25
53.2	30
53.6	36
54.8	50

Takahaschi, 1927

%	f.t.	%	f.t.
50.45	-11 (2+1)	51.90	10
50.93	-5	53.60	30
51.35	0	55.73	50
51.60	+5	58.13	70

Morey and Chen, 1956

f.t.	P Kg
V+L+C	
374	32
400	78
500	161
600	207

Properties of phases

Gerlach, 1859

%	d	%	d
	15°		
0	0.99913		
1	1.00826	27	1.26676
2	.01740	28	.27781
3	.0263	29	.28886
4	.03568	30	.29991
5	.04481	31	.31146
6	.0542	32	.32301
7	.0636	33	.33456
8	.07390	34	.34611
9	.08242	35	.35767
10	.09183	36	.36962
11	.10162	37	.38158
12	.11141	38	.39354
13	.12121	39	.40550
14	.13100	40	.41747
15	.14079	41	.42980
16	.15099	42	.44212
17	.16120	43	.44447
18	.17140	44	.46680
19	.18162	45	.47912
20	.19182	46	.49184
21	.20239	47	.50457
22	.21296	48	.51730
23	.22352	49	.53002
24	.23409	50	.54273
25	.24466	51	.55592
26	.25571	52	.56911

Kohlrausch, 1879

%	d	%	d
	15°		
5	1.0449	30	1.3002
10	.0919	40	.4170
20	.1920	50	.5428

Kuschel, 1881

%	t	d
0.20	23.9	1.0002
0.50	23.6	1.0029
0.50	24.1	1.0023
1.02	21.6	1.0077
1.02	20.42	1.0078
2.94	24.4	1.0244
4.99	22.5	1.0430
10.04	22.1	1.0903
19.75	24.9	1.1833
19.75	20.5	1.1850
39.424	26.0	1.4025

Volkmann, 1882

%	d	%	d
	15°		
4.45	1.0400	27.05	1.2674
8.90	1.0816	34.96	1.3575
16.64	1.1576		
	15-16°		
6.19	1.0562	32.19	1.3254
13.12	1.1227	39.74	.14145
23.74	1.2314		

Slotte, 1883

%	d
20°	
16.31	1.1509
28.66	.2814
39.04	.4026

Moore, 1895-96

N	d	N	d
	18°		
0.00	0.9987	0.9456	1.1100
.273	1.0340	1.974	1.2183
.4788	1.0577		

Pann, 1901

%	d 10°	d 18°	d 30°
0	0.999727	0.998622	0.995673
23.74	1.23245	1.22856	1.22081
39.41	.41322	.40792	.40073
51.15	.56250	.55587	-
53.6	-	-	.56749

Cheneveau, 1907

%	d 16.2°
0	0.9989
34.37	1.3168

Lunge and Berl, 1907

%	d 15°	d 25°	%	d 15°	d 25°
4.9	1.045	1.042	34.8	1.357	1.353
9.8	.091	.088	40.5	.424	.420
15.0	.142	.138	45.2	.483	.483
19.7	.190	.186	50.1	.546	.546
25.5	.252	.248	51.3	.563	.563
29.6	.297	.293			

Wasastjerna, 1920

N (20°)	d (20°)	N (25°)	d (25°)
0.0000	0.99823	0.0000	0.99707
.2117	1.02390	.2114	1.02247
.4300	.04959	.4293	.04799
.6054	.06966	.6044	.06792
.8590	.09799	.8575	.09609
.9822	.11147	.9804	.10949
1.2862	.14416	1.2838	.14204
.5213	.16889	.5185	.16670
.7114	.18856	.7082	.18631
.8765	.20538	.8729	.20307
2.1336	.23120	2.1294	.22878
.2526	.24300	.2482	.24058
.5535	.27247	.5486	.27002
.8009	.29622	.7954	.29367
.9702	.31224	.9644	.30969
3.1768	.33159	3.1706	.32898
.3839	.35069	.3774	.34809

Rehbinder, 1926

%	d (20°)	%	d (20°)
50.0	1.543	24.0	1.299
45.0	.478	20.0	.191
40.0	.416	10.0	.092
35.0	.358	0.0	0.9982

A.E.Hill and D.G.Hill, 1927

%	t	d
52.74	25	1.555
53.66	36	1.557

Takahaschi, 1927

%	t	d
50.45	-11	1.533
50.93	- 5	1.5445
51.35	0	1.546
51.60	+ 5	1.5475
51.90	+10	1.549
53.60	+30	1.547
55.73	+50	1.570
58.13	+70	1.590

Okazaki, 1933

%	d (28°)	%	d (28°)
10.26	1.0898	38.97	.3979
20.45	.1909	47.27	.5005
29.24	.2846		

Hitchcock and Mc Ilkenny, 1935

N(20°)	d 20°	N(30°)	d 30°	N(40°)	d 40°
0.9479	1.0526	0.9447	1.0490	0.9417	1.0453
1.820	.1003	1.813	.0964	1.806	.0920
2.780	.1514	2.768	.1468	2.717	.1422
3.616	.1945	3.602	.1899	3.587	.1850
4.421	.2349	4.403	.2299	4.385	.2248
5.268	.2764	5.246	.2717	5.225	.2663
6.097	.3159	6.073	.3108	6.049	.3054
7.112	.3631	7.084	.3578	7.057	.3525

Guillaume, 1946

%	d
20°	
9.28	1.0855
46.3	1.4952

Slotte, 1883

%	η 20°	η 30°	η 40°
16.31	1519	1230	1018
28.66	2412	1939	1597
39.04	4160	3261	2632

Moore, 1895-96

N	η	N	η
18°			
0.00	1058	0.9456	1386
.25	1121	1.0	1419
.273	1127	1.974	2047
.4788	1198	2.0	2063
.5	1204		

Hitchcock and Mc Ilkenny, 1935

N(20°)	η	N(30°)	η	N(40°)	η
	20°		30°		40°
0.9479	1145.9	0.9447	925.5	0.9417	765.9
1.820	1305.1	1.813	1059.4	1.806	877.1
2.780	1518.4	2.768	1233.5	2.717	1023.3
3.616	1763.8	3.602	1425.7	3.587	1178.0
4.421	2037.2	4.403	1641.0	4.385	1353.9
5.268	2367.7	5.246	1932.1	5.225	1592.6
6.097	2792.5	6.073	2235.5	6.049	1832.4
7.112	3442.2	7.084	2737.6	7.057	2218.0

Volkmann, 1882

%	σ	%	σ
15°			
34.96	87.4	8.90	75.8
27.05	82.0	4.45	74.6
16.64	77.9		
15-16°			
39.74	90.8	13.12	76.7
32.19	85.1	6.19	75.0
23.74	80.5		

Pann, 1901

%	σ 10°	σ 30°
0	74.01	71.01
23.74	81.06	78.07
39.41	92.06	89.23
51.15	108.45	-
53.6	-	107.60

Forch, 1905

N	t	σ
1.517	16.6	78.55
1.898	17.6	79.02
3.585	17.7	82.19
4.483	18.5	84.23
6.291	17.5	88.96

Rehbinder, 1926

%	σ 10°	σ 20°	σ 30°
50.0	-	103.3	101.3
45.0	98.3	96.6	94.9
40.0	92.8	91.4	90.0
35.0	88.6	87.3	86.0
24.0	85.0	83.8	82.5
20.0	79.7	78.6	77.4
10.0	76.5	75.1	73.8
0.0	74.30	72.75	71.17

Walter, 1889

%	n_D
15°	
0	1.3334
1.2	.3355
5.1	.3421
21.1	.3698
50.4	.4180

Jones, 1904; Jones and Getman, 1904

N	n_D	N	n_D
	0°		
0.05	1.32632	0.20	1.33395
.06	.32758	.40	.34553
.10	.32980	1.00	.36219

Chéneveau, 1907

%	t	n C	n D	n Tl	n F	n G'
34.37	16.2	1.38376	1.38594	1.38790	1.39070	1.39468

Wasastjerna, 1920

N(20°)	n H_α	n D	n H_β
		20°	
0.0000	1.33151	1.33330	1.33747
.2117	.33586	.33824	.34248
.4300	.34115	.34307	.34737
.6054	.34483	.34677	.35111
.8590	.34988	.35187	.35619
.9822	.35227	.35427	.35871
1.2862	.35792	.35993	.36450
.5213	.36210	.36416	.36877
.7114	.36535	.36741	.37213
.8765	.36812	.37020	.37493
2.1336	.37228	.37444	.37921
.2526	.37413	.37623	.38109
.5535	.37873	.38090	.38574
.8009	.38235	.38446	.38947
.9702	.38475	.38692	.39193
3.1768	.38761	.38979	.39484
.3839	.39040	.39261	.39765

N(25°)	n H_α	n D	n H_β
		25°	
0.0000	1.33108	1.33290	1.33700
.2114	.33586	.33772	.34191
.4293	.34057	.34246	.34674
.6044	.34417	.34612	.35045
.8575	.34921	.35124	.35561
.9804	.35152	.35354	.35801
1.2838	.35714	.35918	.36377
.5185	.36132	.36416	.36802
.7082	.36458	.36741	.37134
.8729	.36735	.37020	.37415
2.1294	.37154	.37444	.37844
.2482	.37341	.37623	.38029
.5486	.37805	.38090	.38507
.7954	.38164	.38446	.38869
.9644	.38406	.38692	.39118
3.1706	.38689	.38979	.39407
.3774	.38986	.39261	.39690

Jahn, 1891

24.304 g/100 cc 20° (α) magn. = 0.46222

Guillaume, 1946

%	t	n_{5780}	*(α)magn.10^6 (5780Å)
9.28	20	1.3497	1.3497
46.3	19	.4134	1.4134

* in radians, gauss, centim.

Okazaki, 1933

%	Verdet's constant (3441 Å)
28°	
10.25	0.04515
20.45	.04603
29.24	.04771
38.97	.04851
47.27	.04916

Kohlrausch, 1879

%	κ
18°	
5	559
10	1034
20	1800
30	2213
40	2159
50	1462

Jones, 1904 and Jones and Getman, 1904

N	λ
0°	
0.05	57.02
0.10	52.17
0.20	47.77
0.40	43.52
1.00	37.12
2.00	30.78

Next Page

Kuschel, 1881

%	t	transport number (CO_3)
0.20	23.9	0.302
0.50	23.0	.372
1.02	20.42	.377
2.94	24.4	.434
4.99	22.5	.437
10.04	22.1	.416
19.75	20.5	.404
39.42	26.0	.341

Marignac, 1876

%	U
21-52°	
43.45	0.6248
33.87	0.6831
23.51	0.7596
13.43	0.8509
7.13	0.9157
3.70	0.9543
0	1.0000

Water + Potassium sesquicarbonate ($K_2C_2O_5$)

Poggiale, 1843

%	f.t.	%	f.t.
27.67	0	40.08	60
30.26	10	41.66	70
32.44	20	43.24	80
34.47	30	44.70	90
36.36	40	46.10	100
38.30	50		

Water (H_2O) + Potassium acid carbonate ($KHCO_3$)

Poggiale, 1843

%	f.t.	%	f.t.
14.93	0	23.15	40
17.17	10	25.01	50
19.32	20	26.61	60
21.29	30	28.35	70

Mulder, 1866

%	f.t.	%	f.t.
19.32	20	25.02	50
21.29	30	26.61	60
23.15	40	28.35	70

Dibbits, 1874

%	f.t.	%	f.t.	%	f.t
18.34	0	25.26	21	31.76	42
18.67	1	25.57	22	32.06	43
18.99	2	25.87	23	32.38	44
19.35	3	26.20	24	32.71	45
19.68	4	26.52	25	33.02	46
20.00	5	26.82	26	33.33	47
20.35	6	27.15	27	33.64	48
20.69	7	27.43	28	33.95	49
21.00	8	27.75	29	34.24	50
21.35	9	28.06	30	34.60	51
21.62	10	28.37	31	34.91	52
22.00	11	28.67	32	35.23	53
22.33	12	28.98	33	35.54	54
22.56	13	29.30	34	35.86	55
22.99	14	29.60	35	36.19	56
23.32	15	29.92	36	36.51	57
23.64	16	30.21	37	36.83	58
23.90	17	30.53	38	37.15	59
24.30	18	30.84	39	37.50	60
24.61	19	31.15	40		
24.92	20	31.46	41		

E.Hill and D.G.Hill, 1927

%	f.t.	d
26.55	25	1.180
29.9	36	-

Previous Page

Takahaschi, 1927

%	f.t.	d	%	f.t.	d
18.41	0	1.1329	32.24	40	1.2196
21.53	10	.1544	36.04	50	.2439
25.23	20.5	.1772	39.65	60	.3711
28.52	30	.2004	43.37	70	.3005

Kohlrausch, 1879

%	d	κ
	18°	
5	1.0328	370
10	1.0674	686

Moore, 1895-96

N	d	η	N	d	η
		18°			
0.0	0.9987	1058	1.0	1.0581	1186
.25	1.0146	1091	1.98	1.1149	1360
.495	1.0298	1124	2.0	-	1365
.5	-	1127			

Hrynakowski, 1927

t	d	η	σ
	sat.sol.		
50.50	1.2791	1311	72.84

Water (H_2O) + Potassium sulfite (K_2SO_3)

Kremers, 1856

sat.sol. b.t.= 103°

Röntgen and Schneider, 1886

%	d
18°	
8.69	1.0778
16.94	1.1577

%	t	uncorrected π
8.69	18.03	0.801
16.94	17.81	0.642

Traube, 1895

%	d
15°	
0	0.9991
4.70	1.0382
15.68	.1372
24.99	.2298

Water (H_2O) + Potassium acid sulfite ($KHSO_3$)

Platt and Hudson, 1926

%	f.t.	%	f.t.
30.89	14	45.95	69
32.89	20	47.20	73
35.34	31	51.7	83
38.60	40	52.2	90
40.00	50	53.6	100
43.36	60		

Water (H_2O) + Potassium metasilicate (K_2SiO_3)

Morey, 1917

mol%	P		f.t.
27.3	-	(1+1)	200
35.2	4.3		285
44.5	6.1		360
46.3	7.4	(5+1)	380
47.0	6.2	E	370
48.8	10.7		420
58.4	10.8		520
63.5	-		600
65.5	9.0		610
85.8	1.0		942
100	0		976

Traube, 1895

%	d	%	d
	20°		
0	0.99823	14.193	1.13461
5.269	1.04659	17.663	1.17067
9.633	1.08862	27.544	1.28277

Ukihashi, 1956 (fig.)

%	κ	%	κ
0	0	19.7	1300
3.95	600	25.6	1250
6.90	900	27.6	1150
8.86	1100	31.5	800
16.75	1300		

Water (H_2O) + Potassium disilicate ($K_2Si_2O_5$)

Morey, 1917

mol%	f.t.	P	
24.0	285	29.5	(1+1)
33.7	360	38.6	
35.0	380	39.0	
37.2	410	36.0	E
38.1	420	37.7	B
41.6	480	54.1	
43.9	500	66.1	
44.5	520	71.3	
48.0	600	70.9	
93.0	1034	1.0	
100	1041	0	

Water (H_2O) + Potassium selenite (K_2SeO_3)

Traube, 1895

%	d	%	d
	15°		
0	0.9991	8.91	1.0747
3.46	1.0285	18.87	1.1748
3.53	1.0290	19.14	1.1748 ?

Water (H_2O) + Potassium tellurite (K_2TeO_3)

Traube, 1895

%	d
15°	
0	0.9991
4.77	1.0402
11.62	1.1047
20.88	1.2036
23.57	1.2340

Water + Potassium acid selenite ($KH_3Se_2O_6$)

Janitzki, 1932

%	f.t.	%	f.t.
46.52	-6.9	75.71	31.0
53.57	0.0	80.30	40.2
63.20	+11.9	85.55	50.8
68.65	20.3	89.65	59.4

Water + Potassium thiocarbonate (K_2CS_3)

Delachanal, 1877

%	d	%	d
	15°		
0	0.999	28.5	1.208
1.1	1.006	29.8	.218
2.1	.013	31.1	.228
3.1	.021	32.3	.239
4.2	.028	33.6	.250
5.2	.035	35.0	.261
6.3	.043	36.5	.272
7.4	.051	37.8	.283
8.4	.059	39.2	.295
9.6	.066	40.7	.307
10.7	.074	42.0	.319
11.7	.082	43.5	.331
12.8	.090	44.8	.344
13.9	.099	46.2	.356
15.0	.107	47.5	.369
16.1	.115	48.9	.382
17.3	.124	50.4	.396
18.5	.133	51.8	.409
19.6	.142	53.3	.423
20.8	.151	54.9	.437
22.0	.160	56.4	.467
23.2	.170	57.8	.482
24.4	.179	60.8	.497
25.8	.189	62.3	.513
27.1	.198	63.7	.529
		65.1	.545
		66.5	.562
		68.0	.579

Water (H_2O) + Potassium sulfate (K_2SO_4)

Heterogeneous equilibria

Tammann, 1885

t	p		
	0%	10.65%	12.80%
64.50	183.5	180.1	179.4
66.83	203.5	199.6	198.2
70.95	243.6	238.6	238.2
73.82	275.3	269.9	269.5
75.74	298.4	292.4	291.2
79.21	344.3	337.5	336.8
81.49	377.5	369.8	367.5
84.91	432.3	423.0	420.0
88.05	488.2	477.8	474.9
91.37	553.9	542.7	541.4
93.77	605.9	593.1	589.6
96.80	677.1	664.1	659.3
100.43	771.6	753.3	749.7

%	p	%	p
	100°		
5.93	749.9	15.52	731.8
9.06	744.6	17.50	727.9
10.58	741.6		

Pohle, 1927

t	p	t	p
	sat.sol.		
30.0	31.0	70.0	223.0
35.0	40.0	75.0	275.0
40.0	54.0	80.0	336.0
45.0	70.0	85.0	410.0
50.0	90.0	90.0	493.0
55.0	115.0	95.0	615.0
60.0	144.0	100.0	708.0
65.0	180.0		

Leopold and Johnston, 1927

m	t	p	
		0%	sat.sol.
0.6256	18.90	16.37	15.91
.6620	22.25	20.13	19.55
.6884	24.73	23.38	22.60
.6974	25.58	24.59	23.89
.7600	31.55	34.77	33.61
.8052	35.96	44.47	42.90
.8825	43.42	66.23	63.74
.9333	48.41	85.46	82.11
.972	52.30	103.60	99.13

Adams and Merz, 1929

t	p	t	p
	sat.sol.		
10	9.06	30	30.68
15	12.78	40	53.04
20	17.30	50	88.57
25	23.56		

Adams, 1932

%	P Kg	%	P Kg	
		25°		
10.76	1	15.9	6012	
13.8	1002	15.4	7014	
15.4	2004	14.9	8016	sat.sol.
16.1	3006	14.4	9018	
16.3	4008	13.8	10020	
16.2	5010	13.2	10022	

Foote, Saxton and Dixon, 1932

$$\lg P = -\frac{2332.5}{T} + 9.1881$$

Morey and Chen, 1956

f.t.	P Kg
sat.sol.	
374	255
400	537
500	868

Stokes, 1948

m	osmotic coefficient	activity coefficient
	25°	
0.1	0.779	0.441
0.2	.742	.360
0.3	.721	.316
0.4	.703	.286
0.5	.691	.264
0.6	.679	.246
0.7	.670	.232

Kremers, 1856

Sat.sol. b.t.= 103°

Gerlach, 1886

%	b.t.	%	b.t.
0	100	18.10	101.5
6.54	100.5	23.08	102
12.66	101	24.01	102.1

Baroni, 1893

%	b.t.	%	b.t.
1.52	100.145	7.90	101.523
3.37	100.229	11.06	101.740
5.56	100.371	15.01	101.930

Buchanan, 1899

%	b.t.	%	b.t.	%	b.t.
767.65 mm		616.37 mm		552.27 mm	
satd	101.70	satd	95.70	satd	92.53
19.28	101.66	18.37	95.60	18.26	92.51
18.80	101.61	16.00	95.40	12.39	92.11
17.17	101.49	14.34	95.30	0.00	91.30
14.48	101.28	13.14	95.20		
0.00	100.28	11.71	95.10		
		10.12	95.00		
		8.23	94.90		
		0.0	94.42		

Barnes and Mc Intosh, 1909

%	b.t.
4.525	100.28
8.010	100.53
11.857	100.86
15.102	101.08
18.842	101.29

Berkeley and Applebey, 1911

Sat.sol. b.t.= 101.392°

Mulder, 1866

%	f.t.	%	f.t.
7.80	0	14.16	53
9.42	16.75	15.90	65
10.03	23	21.10	102.25 b.t.
11.42	31		

Nordenskjöld, 1869

%	f.t.	%	f.t.
7.24	0.0	13.79	47.0
9.34	15.65	16.87	70.2
11.34	28.1	19.28	98.0

Tilden and Shenstone, 1884

%	f.t.	%	f.t.
8.89	16	14.81	59
9.34	20	18.86	98
11.18	28	20.95	120
11.72	36	22.36	143
12.44	39	24.75	170

Andreae, 1884

%	f.t.	%	f.t.
6.84	0	12.85	40
7.49	4	14.16	50
8.59	11	15.37	60
10.00	20	16.50	70
11.48	30		

Etard, 1894

%	f.t.	%	f.t.
10.1	21	22.8	152
10.3	23	24.5	175
14.5	60	23.8	195
19.1	99	24.6	220
21.2	130		

Berkeley, 1904

%	f.t.	%	f.t.
6.95	0.40	15.26	58.95
9.39	15.70	17.11	74.85
11.76	31.45	18.57	89.70
13.42	42.75	19.48	101.1

Chernai, 1912

%	f.t.	%	f.t.
6.84	0	12.86	40
8.44	10	14.16	50
10.00	20	15.37	60
11.48	30	16.49	70

Le Blanc and Schmandt, 1911

%	f.t.
7.82	4.78
11.43	30.05
14.77	54.20
16.40	68.90

Lattey, 1923

%	f.t.
9.15	14.5
11.70	30.2
14.53	51.9

Cornec and Hering, 1925

%	f.t.	%	f.t.
6.75	- 1.52	14.19	50
6.83	0	17.09	75
10.82	25	19.42	100

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
3.9	357	26.4	289
8.36	344	26.4	274
12.3	337	25.9	245
16.2	327	25.8	215
23.3	310	25.7	208
25.8	303	25.6	193
26.0	292	25.4	179

Bergman and Sholokhovich, 1942

%	f.t.	%	f.t.
1	-0.5	5	-1.5
2	-0.7	6	+8.2
3	-0.9	7	20.4
4	-1.2	8	24.6

E: 6.5% -1.5°

Chernogorenko, 1956

E_1 : -1.9 E_2 : -1.5 (1+1)

Properties of phases.

Gerlach, 1859 and 1889

%	d (1859)	d (1889)
	15°	
1	1.00732	1.00720
2	.01547	-
3	.02361	1.02358
4	.03187	-
5	.04014	1.04000
6	.04856	-
7	.05698	1.05684
8	.06551	-
9	.07405	1.07405
9.92	-	1.08210

Berkeley, 1904

%	t	d
	sat. sol.	
6.95	0.40	1.0589
9.39	15.70	1.0770
11.76	31.45	1.0921
13.42	42.75	1.1010
15.26	58.95	1.1086
17.11	74.85	1.1157
18.57	89.70	1.1194
19.48	101.1	1.1207

Chéneveau, 1907

%	d	%	d
	15°		
0	0.9991	5.68	1.0464
1.26	1.0105	6.83	.0563
2.49	.0203	7.96	.0658
4.11	.0334		

d^{19}= 1.0637 8.55%

Chernai, 1912

%	t	d
6.84	0	1.058
8.44	10	1.069
10.00	20	1.081
11.48	30	1.089
12.86	40	1.097
14.16	50	1.106
15.37	60	1.114
16.49	70	1.121

Manchot, Jarhstorfer and Zepter, 1924

c	d
25°	
10.440	1.0762
10.355	1.0753

Cornec and Hering, 1925

%	t	d
10.82	25	1.086
14.19	50	1.104
17.09	75	1.116
19.42	100	1.119

Hrynakowski, 1927

Sat. sol. 50.15° d = 1.1112

Jones and Colvin, 1940

M	d 25°	d 0°
0.0005	0.99715	0.99993
.001	.99723	1.00001
.002	.99736	.00015
.005	.99779	.00061
.01	.99850	.00137
.02	.99990	.00284
.05	1.00407	.00724
.1	.01094	.01462
.2	.02443	.02876
.3	.03774	.04270
.5	.06386	-

Hrynakowski, 1927

Sat. sol. 50.15° η = 740

Jones and Colvin, 1940

M	η(water=1) 25°	0°
0.0005	1.00039	1.00028
.001	.00062	.00040
.002	.00103	.00061
.005	.00198	.00101
.01	.00334	.00157
.02	.00586	.00241
.05	.01281	.00441
.1	.02380	.00795
.2	.04535	.01614
.3	.06760	.03040
.5	.11470	-

Volkmann, 1882

%	σ		%	σ
4.49	74.6	15-16°	9.01	75.2

Hrynakowski, 1927

Sat. sol. 50.1° σ = 67.73

Borgesius, 1895

%	t	D n_D (sol.-aq.)
5.12	18.8	0.000684
17.59	18.6	0.002661

Chèneveau, 1907

%	t	n C	n D	n Tl	n F	n G'
8.55	29.0	1.34134	1.34322	1.34518	1.34739	1.35064

Kohlrausch, 1879

%	κ		%	κ
5	456	18°	10	856

Scholz, 1892

N	Q diss (cal/gr)	N	Q diss (cal/gr)
0.0625	-48.57	0.5	-45.14
0.125	-48.26	0.977	-42:09
0.25	-47.06		

Mondain-Monval, 1923

Q diss. 0.8 - 0.9 mol %
0° -8020 18° -6400

Q dil. (sat. sol. - 0.8 mol %)
0° -870 18° -670

Q mix (large amount of sat. sol. + 1 mole aqua)
0° -4 18° -6

Q diss. (large amount of sat. sol. + 1 mole salt)
0° -540 18° -570

Q diss. limit
0° -6750 18° -5900

Water + Potassium acid sulfate ($KHSO_4$)

Tammann, 1885

%	p	%	p
	100°		
10.04	742.0	35.40	673.0
17.01	726.8	43.78	634.8
23.96	708.7	46.33	623.1
34.87	675.1	53.63	580.4

Kohlrausch, 1879

%	d	%	d
	15°		
5	1.0354	20	1.1516
10	.0726	25	.1920
15	.1116	27	.2110

Grotrian, 1879

t	d	η	t	d	η
		19.63%			
15.49	1.1500	1506.5	29.19	1.1422	1100.1
16.15	.1496	1488.6	38.58	.1370	916.8
29.07	.1423	1105.9	38.99	.1368	910.6

Moore, 1895 - 1896

N	d	η (water=1)
0.0	0.9987	1.0000
0.5	1.0439	1.075
1.0	1.0866	1.149
2.0	1.1712	1.352

Randall and Taylor, 1941

N	d
25°	
2.4310	1.1922

Hrynakowski, 1927

Sat. sol. t = 50.70° d = 1.4794
η = 2082 σ = 72.48

Kohlrausch, 1879

%	κ	%	κ
18°			
5	817	20	2762
10	1524	25	3246
15	2173	27	3409

Gillespie and Wasif, 1953

N	κ	N	κ
25°			
0.0568	125.6	0.8686	583.4
0.1251	174.0	1.0278	636.0
0.1981	231.5	1.1568	669.5
0.2532	217.0	1.3336	711.2
0.3602	338.0	1.4074	722.4
0.4934	415.6	1.6148	759.3
0.6242	478.3		

Randall and Taylor, 1941

N	U	N	U
25°			
0.0498	0.99720	0.5547	1.00595
0.1017	0.99725	1.0151	1.01884
0.2483	0.99919	2.2442	1.06272

Water (H_2O) + Potassium selenate (K_2SeO_4)

Etard, 1894

%	f.t.	%	f.t.
51.5	-20	52.6	+18
51.7	- 5	54.9	+97
52.0	+ 5		

Friend, 1929

%	f.t.	%	f.t.
53.57	0.0	54.39	51.2
53.59	11.6	54.49	52.4
53.63	14.0	54.69	59.0
53.66	19.4	54.97	68.5
53.81	26.8	54.96	71.0
54.04	30.4	55.31	80.6
54.08	32.6	55.50	81.2
54.17	38.0	55.85	89.6
54.33	47.4	56.16	97.6

Traube, 1895

%	d	%	d
15°			
0	0.99913	8.709	1.07281
1.904	1.01484	12.702	1.10945
2.776	1.02205	13.682	1.11846
3.818	1.03078		

Tutton, 1897

%	d	%	d
20°			
35.76	1.3591	50.00	1.5590
41.79	1.4385		

Water (H_2O) + Potassium tellurate (K_2TeO_4)

Rosenheim and Weinheber, 1911

%	f.t.
8.11	0
21.59	20
33.52	30

Water + Potassium thiosulfate ($K_2S_2O_3$)

Tammann, 1885

%	p	%	p
100°			
5.78	751.0	36.81	675.4
9.95	744.4	43.86	642.5
17.46	730.7	49.43	610.1
24.30	715.1	58.56	540.2
28.34	704.6	66.26	472.4
31.14	696.7		

Jo, 1911

%	f.t.	%	f.t.	%	f.t.
49.01	0	67.41	43.0	71.87	70.0
60.03	17.2	67.70	46.0	72.71	75.0
61.03	20.0	68.28	50.0	73.81	77.0
61.90	24.3	69.42	54.5	74.26	78.0
62.61	26.6	69.85	56.0	74.52	79.5
63.17	28.1	70.28	58.4	74.69	82.2
64.08	31.0	70.62	61.0	74.93	85.2
65.09	33.5	71.18	66.0	75.52	89.0
66.93	35.0				

Equilibrium temp.

salt - (1+3)	78.3	(1+3) - (1+1)	56.1
(1+1) - (5+3)	35.0	(5+3) - (2+1)	0 - 17

Water + Potassium pyrosulfate ($K_2S_2O_7$)

Kremers, 1854

%	f.t.	%	f.t.
25.32	0	38.61	40
32.43	20	59.95	100

Traube, 1895

%	d	%	d
	15°		
0	0.99823	7.725	1.05857
2.175	1.01578	12.968	1.09754
3.164	1.02325	22.207	1.17758

Water + Potassium dithionate ($K_2S_2O_6$)

Tammann, 1885

%	p	%	p
	100°		
10.65	747.8	29.43	718.5
18.71	736.7	32.56	713.2
23.87	729.4		

Kurtenacher and Laszlo, 1938

2nd comp.	formula	% 0°	% 20°
Potassium trithionate	$K_2S_3O_6$	8.1	18.4
Potassium tetrathionate	$K_2S_4O_6$	12.6	23.2
Potassium pentathionate	$K_2S_5O_6$	15.5	24.8

Water + Potassium pyroselenate ($K_2Se_2O_5$)

Janitzki, 1932

%	f.t.	%	f.t.
73.52 (1+1)	-20.6	79.18	24.8
74.29	-10.5	79.21	27.2
75.87	+0.2	79.23	29.4
77.21	12.8	79.41	31.7
78.18	18.9	79.99	39.8
78.50	20.6	80.39	50.4
78.70	23.0	81.69	59.8
79.31	25.2	81.55	65.4
79.63	27.9	82.42	69.6
80.04	30.6	82.18	72.6
79.01 B	30.8	83.72	90.8
79.18 E	24.0	84.43	102.8

Water + Potassium formate ($KCHO_2$)

Tammann, 1885

%	p	%	p
	100°		
7.60	737.0	37.89	567.1
10.35	726.9	38.79	559.8
18.18	692.0	50.12	461.7
22.02	673.4	56.44	408.1
27.66	639.9	63.16	343.6
32.50	607.7	66.28	313.3

Groschuff, 1903

%	f.t.	%	f.t.
72.8	-20	92.0	120
76.8	+18	96.0	140
80.7	50	100.0	157
86.8	90		

Sidgwick and Gentle, 1922

%	f.t.	%	f.t.
5.71	-2.62	25.03	-15.82
12.95	-6.72	26.04	-17.39
18.42	-10.42	100	+157

de Garcia, 1920

N	d	n_D	N	d	n_D
		21°			
4	1.176300	1.3601	0.25	1.011249	1.3349
2	.090912	.3491	.125	.005318	.3337
1	.046098	.3409	.0675	.002278	.3331
0.5	.024312	.3371			

Water + Potassium acid formate ($KC_2H_3O_4$)

Groschuff, 1903

%	f.t.
60.4	0
69.8	25
79.2	50
90.7	80

Water + Potassium acetate ($KC_2H_3O_2$)

Heterogeneous equilibria

Tammann, 1885

%	p	%	p
	100°		
7.66	739.0	45.27	496.4
12.41	722.6	51.51	432.8
21.40	681.4	59.82	344.5
27.45	646.4	66.08	281.0
33.34	604.2	73.54	203.4
39.19	553.2		

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1263	0.1230	1.442	1.245
.1621	.1575	.741	.480
.1872	.1821	.793	.518
.2118	.2037	2.015	.685
.3397	.3233	.060	.716
.3593	.3417	.244	.855
.5185	.4830	.365	.947
.5922	.5496	.479	2.035
.8090	.7326	.896	.336
.9231	.8295	3.113	.487
1.017	.9012	.568	.810
.103	.9766	.630	.834
.232	1.085	.983	3.099
.241	.088	4.480	.458
.435	.236	.516	.475

m_1= molality of potassium chloride

m_2= molality of potassium acetate

Robinson and Stokes, 1949

m	osmotic coefficient
25°	
0.1	0.943
.2	.944
.3	.951
.4	.958
.5	.968
.6	.977
.7	.987
.8	.997
.9	1.007
1.0	.017
.2	.038
.4	.060
.6	.081
.8	.103
2.0	.123
2.5	.177
3.0	.228
3.5	.274

Legrand, 1835

%	b.t.	%	b.t.
0.0	100	64.30	126
9.50	101	65.27	127
16.67	102	66.22	128
22.24	103	67.15	129
26.69	104	68.04	130
30.26	105	69.75	132
33.24	106	71.31	134
35.82	107	72.78	136
38.12	108	74.19	138
40.26	109	75.51	140
42.30	110	76.79	142
44.22	111	78.02	144
46.04	112	79.19	146
47.75	113	80.31	148
49.39	114	81.20	150
50.96	115	82.38	152
52.45	116	83.33	154
53.87	117	84.23	156
55.23	118	85.07	158
56.54	119	85.86	160
57.79	120	86.61	162
58.98	121	89.34	164
60.13	122	87.96	166
61.22	123	88.57	168
62.27	124	88.87	169
63.31	125		

Gerlach, 1886

%	b.t.	%	b.t.
0	100	68.80	131
5.66	101	69.65	132
10.71	102	70.45	133
15.25	103	71.01	134
19.88	104	71.95	135
23.67	105	72.67	136
27.28	106	73.40	137
30.55	107	74.14	138
33.55	108	74.84	139
36.31	109	75.55	140
38.84	110	76.25	141
41.18	111	76.93	142
43.50	112	77.59	143
45.65	113	78.20	144
47.64	114	78.79	145
49.49	115	79.40	146
51.22	116	80.00	147
52.82	117	80.56	148
54.34	118	81.09	149
55.85	119	81.63	150
57.27	120	82.16	151
58.58	121	82.65	152
59.84	122	83.14	153
61.01	123	83.59	154
62.12	124	84.03	155
63.17	125	84.44	156
64.22	126	84.67	157
65.21	127	85.19	158
66.16	128	85.55	159
67.04	129	85.89	160
57.94	130	86.22	161

Abe, 1911

%	f.t.	%	f.t.
68.42	0.1	76.63	41
69.12	5	76.71	42
70.01	10	76.79	43.5
70.86	15	76.85	45
71.88	20	77.11	50
72.44	23	77.44	55
72.93	25	77.78	60
73.95	30	78.09	65
75.11	35	78.49	70
75.85	38	79.19	80
76.01	39	79.85	90
76.37	40	80.23	96
(3+1)	(2+1)		

Guthrie, 1876

%	f.t.	%	f.t.
5	-2	20	-11.2
10	-4.6	25	-16
15	-7.4	30	-22.5

Sidgwick and Gentle, 1922

%	f.t.	%	f.t.
5.00	-2.15	20.53	-12.32
7.42	-3.36	25.50	-17.92
17.04	-9.32	100	+292

Fricke and Schützdeller, 1924

mol%	f.t.	mol%	f.t.
1.887	-8.143	0.180	-0.673
1.476	-6.119	.072	-0.266
0.938	-3.540	.028	-0.103
.364	-1.293		

Properties of phases.

Gerlach, 1869

%	d	%	d
	17.5°		
0	0.9987	40	1.2090
10	1.0476	50	.2669
20	.0991	60	.3268
30	.1530	sat.sol.	.43

Hager, 1876

%	d	%	d
	17.5°		
10	1.04861	40	.2116
20	.1003	50	.2710
30	.1548		

Kohlrausch, 1879

%	d	%	d
	15°		
5	1.0228	40	1.2028
10	.0466	50	.2590
20	.0960	60	.3152
30	.1484	70	.3714

Gerlach, 1887

%	d	%	d
	17.5°		
1	1.0036	31	1.1585
2	.0085	32	.1640
3	.0134	33	.1695
4	.0183	34	.1759
5	.0232	35	.1801
6	.0381	36	.1861
7	.0330	37	.1918
8	.0379	38	.1975
9	.0428	39	.2032
10	.0477	40	.2089
11	.0526	41	.2146
12	.0576	42	.2203
13	.0620	43	.2260
14	.0676	44	.2317
15	.0726	45	.2374
16	.0729	46	.2433
17	.0832	47	.2491
18	.0885	48	.2550
19	.0938	49	.2609
20	.0991	50	.2668
21	.1044	51	.2727
22	.1097	52	.2786
23	.1150	53	.2845
24	.1203	54	.2904
25	.1256	55	.3063
26	.1311	56	.3024
27	.1365	57	.3085
28	.1420	58	.3146
29	.1475	59	.3207
30	.1530	60	.3268

Zecchini, 1905

%	t	d
9.599	9.3	1.05091
16.335	6.1	.08643
22.810	9.2	.12165
31.437	6.7	.17201

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.00111	1.0	1.0468
0.1	.00364	2.0	.0913
0.2	.00855	4.0	.1762
0.5	.02278		

Schneider, 1910

N	d	N	d
	18°		
0	0.99812	1.00	1.0468
0.05	1.00112	1.907	1.0872
0.1	1.00365	3.814	1.1698
0.2	1.00858		
0.5	1.0231		

Wasastjerna, 1920

N	d	N	d
18°		25°	
0.0000	0.99862	0.0000	0.99707
0.6346	1.02966	0.6334	1.02774
1.8091	.08268	1.8050	.08021
3.0494	.13718	3.0422	.13448
3.4130	.15244	3.4042	.14947
5.1609	.22424	5.1460	.22070
8.3676	.34403	8.3405	.33967

de Garcia, 1920

N	d	N	d
	18°		
4	1.160905	0.25	1.010835
2	.083022	.12	.005374
1	.042450	.06	.002679
0.5	.021384		

Fricke and Schützdeller, 1924

N	d
25°	
1.785	1.0746
1.383	1.0584
0.892	1.0350
0.357	1.1007
0.1785	1.0026

mol%	d	mol%	d
	13-17°		
1.711	1.0768	0.178	1.0066
1.368	1.0611	.071	.0020
0.892	.0390	.028	0.9999
.357	.0148		

Bury and Parry, 1935

%	d	%	d
	25°		
1.005	1.00204	29.80	1.1536
2.005	.00686	35.26	.1846
3.787	.01575	39.78	.2110
4.870	.02111	41.71	.2220
6.418	.02883	44.67	.2401
8.102	.03733	46.61	.2533
8.938	.04148	49.05	.2655
9.130	.04649	50.59	.2749
11.03	.05233	53.13	.2909
12.84	.06150	54.51	.2992
15.01	.07270	55.68	.3060
18.60	.09181	57.25	.3149
24.86	.1260	60.82	.3373
25.05	.1264	61.56	.3417
29.01	.1491	62.30	.3453

Guillaume, 1946

%	d
20°	
5.27	1.0265
11.0	.0525
19.0	.0954
22.6	.1165
31.5	.1651

Schneider, 1910

N	η (water=1)	N	η (water=1)
18°			
0.05	1.0150	1.00	1.2619
0.1	.0288	1.907	1.5481
0.2	.0523	3.814	2.5059
0.5	.01284		

Fricke and Schützdeller, 1924

N	η
25°	
1.785	1290.2
1.303	1197.2
0.892	1075.7
0.357	967.6
0.1785	930.2

Walter, 1889

%	n_D
15°	
0	1.3334
7.0	.3414
21.5	.3584
44.4	.3862

Zecchini, 1905

%	t	n_D
9.599	9.3	1.34590
16.335	6.1	.35437
22.810	9.2	.36216
31.437	6.7	.37398

Rubien, 1911

N	n_D	N	n_D
18°			
0	1.33327	0.9795	1.34417
0.1005	.33444	1.9405	.35425
.2021	.33557	2.752	.37174
.4964	.33884		

Heydweiller, 1913

N	n_D	N	n_D
18°			
0	1.33327	1.0	1.34439
0.1	.33444	2.0	.35487
0.2	.33555	4.0	.37402
0.5	.33888		

Wasastjerna, 1920

N	n H_α	n D	n H_β
18°			
0.0000	1.33168	1.33348	1.33764
0.6346	.33895	.34084	.34515
1.8091	.35087	.35294	.35749
3.0494	.36312	.36520	.36991
3.4130	.36637	.36851	.37323
5.1609	.38173	.38392	.38897
8.3676	.40556	.40787	.41339
25°			
0.0000	1.33108	1.33291	1.33701
0.6334	.33817	.34008	.34425
1.8050	.34997	.35208	.35652
3.0422	.36216	.36426	.36889
3.4042	.36535	.36740	.37216
5.1460	.38060	.38275	.38775
8.3676	.40436	.40787	.41204

de Garcia, 1920

N	n_D	N	n_D
18°			
4	1.3730	0.25	1.3361
2	.3537	.12	.3346
1	.3439	.06	.3340
0.5	.3389		

Guillaume, 1946

%	$n_{5780\ \text{Å}}$
20°	
5.27	1.3406
11.0	.3464
19.0	.3545
22.6	.3611
31.5	.3718

%	*$(\alpha)_{magn.}10^6$ (5780 Å)
	20°
5.27	3.912
11.0	3.838
19.0	3.725
22.6	4.677
31.5	3.570

* in radians, gauss, centim.

Kohlrausch, 1879

%	κ	%	κ
		18°	
5	346	40	1256
10	623	50	1117
20	1043	60	840
30	1250	70	476

Heydweiller, 1909

N	λ	N	λ
		18°	
0.05	87.7	1.0	63.4
0.1	83.8	2.0	51.4
0.2	79.2	4.0	32.0
0.5	71.6		

Thermal constants.

Marignac, 1876

%	U
20-51°	
52.14	0.6391
30.35	.7728
17.89	.8572
9.82	.9170
5.17	.9550
0	1.0000

Urban, 1932

t	U			
	2.360 M*	0.974M	0.479M	0.249M
10	0.8440	0.9202	0.9528	0.9816
15	.8479	.9240	.9553	.9843
20	.8515	.9266	.9561	.9849
25	.8530	.9271	.9560	.9845
30	.8530	.9270	.9555	.9835
35	.8526	.9266	.9548	.9824
40	.8520	.9259	.9542	.9813

M* = moles acetate in 1000 g solution

Water + Potassium propionate ($KC_3H_5O_2$)

Tammann, 1885

%	p	%	p
		100°	
10.73	733.4	48.43	516.3
19.39	702.4	53.53	454.9
40.19	578.4	59.45	405.6

de Garcia, 1920

N	d	N	d
		20°	
0.06	1.001932	1	1.043826
0.12	.004756	2	.086631
.25	.010334	4	.144252
.5	.021461		

Heydweiller, 1921

N	d	N	d
		18°	
0.2	1.00968	2	1.0922
0.5	.07388	3	.1352
1	.04705	4	.1752

de Garcia, 1920

N	n_D	N	n_D
		20°	
0.06	1.3340	1	1.3464
.12	.3348	2	.3592
.25	.3364	4	.3785
.50	.3396		

Heydweiller, 1921

N	λ	N	λ
		18°	
0.2	73.7	2	42.73
0.5	65.8	3	32.37
1	56.4	4	24.68

Water + Potassium butyrate ($KC_4H_7O_2$)

Tammann, 1885

%	p	%	p
	100°		
13.50	728.6	43.51	588.3
20.17	705.6	46.21	570.0
34.75	639.9	59.16	460.7

de Garcia, 1920

N	d	n_D	N	d	n_D
		20°			
0.06	1.002300	1.3336	1	1.043881	1.3489
.12	.005124	.3350	2	.086350	.3641
.25	.010610	.3371	4	.157910	.3904
.50	.021924	.3411			

Vorländer and Kirchner, 1930-31

d	η (water=1)	B
	20°	
1.232	19.3	-0.64
.234	23.5	-0.85
.245	30.9	-0.99

B = specific mechanical double birefringence

Water + Potassium isobutyrate ($KC_4H_7O_2$)

Tammann, 1885

%	p	%	p
	100°		
20.05	708.3	47.92	559.8
29.86	666.6	55.98	492.1
41.01	605.4	65.67	392.3

de Garcia, 1920

N	d	n_D	N	d	n_D
		25°			
0.06	1.001123	1.3336	1	1.044835	1.3485
.12	.004188	.3346	2	.089080	.3640
.25	.010030	.3366	4	.167829	.3909
.50	.021801	.3408			

Water + Potassium valerate ($KC_5H_9O_2$)

Tammann, 1885

%	p	%	p
	100°		
16.47	726.4	38.02	651.8
25.76	698.7	55.65	549.7
35.28	663.4	70.20	405.0

Water + Potassium isovalerate ($KC_5H_9O_2$)

Vorländer and Kirchner, 1930 - 1931

66% 22.5° d=1.177 η(water=1)=45.5

specific mechanical birefringence = -0.77

Water + Potassium caprylate ($KC_8H_{15}O_2$)

Bunbury and Martin, 1914

%	d	κ	%	d	κ
			90°		
0	0.9578	-	3.61	-	357.8
0.18	-	23.09	8.38	0.9708	757.2
0.90	-	106.1	15.49	.9816	1251
1.80	-	195.2	35.92	-	2165

Davies, 1933

%	U	%	U
	13.5°-16.5°		
1.864	0.9957	7.000	0.9853
2.842	.9930	7.535	.9844
3.927	.9908	7.897	.9825
5.000	.9877	8.936	.9787
5.492	.9857	10.83	.9666
5.940	.9852	14.88	.9305
6.489	.9857	19.86	.8901

Water + Potassium laurate ($KC_{12}H_{23}O_2$)

Bunbury and Martin, 1914

%	d	κ	%	d	κ
		90°			
0	0.9579	-	10.67	0.9645	636
0.24	-	22.45	15.23	-	888
1.18	-	93.41	19.34	0.9685	1134
2.32	-	150.90	32.59	0.9759	1650
4.55	-	266.80			

Bury and Parry, 1935

%	d		%	d	
	25°	35°		25°	35°
0.300	0.99749	0.99442	3.042	1.00021	0.99693
.595	.99789	.99478	4.001	.00105	.99767
.643	-	.99482	4.006	.00110	.99774
.718	-	.99491	5.998	.00289	.99939
.797	0.99813	.99500	8.002	.00470	1.00104
1.000	.99813	.99517	9.000	.00550	.00174
.501	.99881	.99560	11.02	.00740	.00351
.505	.99883	.99563	12.52	.00861	.00458
2.000	.99927	.99604	14.05	.01000	.00587
.001	.99927	.00604	17.03	.01514	.01057
.987	1.00015	.99683	18.21	.01351	.01907
3.001	1.00016	.99684	20.21	.01247	.01814

Water + Potassium myristate ($KC_{14}H_{27}O_2$)

Bunbury and Martin, 1914

%	d	κ	%	d	κ
		90°			
0	0.9579	-	2.59	-	114.7
0.27	-	21.59	5.07	-	240.5
0.53	-	34.88	11.79	0.9618	580.1
1.32	-	65.09	21.14	.9643	1051

Water + Potassium palmitate ($KC_{16}H_{31}O_2$)

Bunbury and Martin, 1914

%	d	κ	%	d	κ
		90°			
0	0.9579	-	5.57	-	202.8
0.29	-	16.51	12.87	0.9593	536.5
0.58	-	25.56	18.16	.9597	763
1.45	-	52.70	22.86	.9604	933
2.86	-	100.70			

Water + Potassium stearate ($KC_{18}H_{35}O_2$)

Bunbury and Martin, 1914

%	d	κ	%	d	κ
		90°			
0	0.9579	-	6.07	-	181.4
0.32	-	14.21	13.91	0.9571	473.9
0.64	-	23.97	19.55	-	658.3
1.58	-	48.34	24.50	-	830
3.12	-	89.80			

Water + Potassium oleate ($KC_{18}H_{33}O_2$)

Dennhardt, 1899

N(18°)	λ					
	0°	10°	18°	25°	40°	60°
0.5	330	434	526	620	816	1072
.4	320	415	507.5	590	791	1045
.3	310	400	486.7	563	743	1007
.2	300	390	470	535	715	960
.15	286	376	450	533	713	954
.1	285	368	445	515	712	930
.05	260	340	420	500	690	933
.03	230	321	383	517	702	960
.02	300	330	380	550	750	1035
.01	-	340	400	570	900	1200
.005	-	-	500	800	1000	1400

Water + Potassium oxalate ($K_2C_2O_4$)

Tammann, 1885

%	p	%	p
	100°		
7.95	745.5	35.15	656.1
15.80	727.8	36.31	650.7
24.13	702.7	42.52	611.4

Lescoeur, 1896

t	p dissoc.	t	p dissoc.
	(2+1)		
50	20	80	69
60	23	90	132
70	41	100	215

Koppel and Cahn, 1908

%	b.t.
7.34	101.05
20.47	101.80
30.31	103.30
40.96	105.09

%	f.t.	%	f.t.
3.020	-0.78	28.50	+30
5.680	-1.49	30.44	40
9.195	-2.50	32.60	50
11.855	-3.22	34.72	60
15.560	-4.38	36.75	70
19.43	-5.88	38.875	80
20.35	0	40.90	90.2
24.17	+10	44.19	106.2
26.675	20		

Hartley, Drugman and al., 1913

%	f.t.	%	f.t.
20.33	0.02	25.85	20.05
21.78	5.05	27.37	25.15
23.08	9.84	28.36	29.98

(1+1)

Colani, 1916

%	f.t.
24.09	15
25.41	19
32.75	50

Fricke and Schützdeller, 1924

m	f.t.	m	f.t.
1.149	-4.553	0.159	-0.691
0.977	-3.887	.107	-0.482
.545	-2.214	.053	-0.250
.304	-1.267		

Voskresenskaya, 1928

%	f.t.	E	%	f.t.	E
6.63	-1.85	-	18.63	-5.7	-6.3
10.47	-3.0	-	18.83	-5.75	-6.4
12.44	-3.6	-	19.19	-5.9	-6.3
14.10	-4.0	-6.3	19.89	-6.3	-6.3
15.79	-4.8	-6.2	20.89	-	-6.4
17.12	-5.05	-6.4	22.64	-	-6.4
17.92	-5.4	-6.4	28.77	-	-6.3

f.t.	%	f.t.	%
0	20.16	50	33.51
10	23.26	60	35.65
16	24.81	70	37.81
20	25.89	80	40.16
25	27.41	90	42.27
30	28.72	100	44.53
40	32.98	107	45.80

Benrath, 1942

%	f.t.	%	f.t.
50	130	68	239
55	150	75	290
60	177	80	330
65	219		

E: 19.92% -6.34°

Franz, 1872

% (1+1)	d	% (1+1)	d
	17.5°		
0	0.9987	13	1.0834
1	1.0054	14	.0898
2	.0121	15	.0963
3	.0188	16	.1029
4	.0255	17	.1095
5	.0324	18	.1161
6	.0388	19	.1227
7	.0452	20	.1292
8	.0515	21	.1357
9	.0579	22	.1433
10	.0642	23	.1489
11	.0706	24	.1555
12	.0770	25	.1623

Kohlrausch, 1879

%	d
18°	
5	1.0367
10	1.0751

Moore, 1895-96

M	d
18°	
0.00	0.9987
0.25	1.0283
0.50	.0571
1.0	.1121
1.5	.1663

Wasastjerna, 1920

M	d	M	d
18°		25°	
0.0000	0.99862	0.0000	0.99707
.0824	1.00887	.0823	1.00721
.2120	.02457	.2116	.02273
.3301	.03869	.3295	.03667
.4227	.04944	.4219	.04740
.6279	.07308	.6266	.07084
.6838	.07941	.6823	.07711
.8449	.09752	.8430	.09519
.9642	.11073	.9620	.10816

de Garcia, 1920

N	d	N	d
17°			
0.06	1.003392	0.5	1.029290
.12	.007050	1	.057602
.25	.014540	2	.112920

Fricke and Schützdeller, 1924

N	d	N	d
25°			
2.158	1.1153	0.598	1.0292
1.854	.0989	.315	.0131
1.058	.0552	.212	.0063

Moore, 1895-96

M	η (water=1)
18°	
0.00	1.000
.25	.049
.50	.103
1.0	.232
1.5	.389

Fricke and Schützdeller, 1924

N	η	N	η
25°			
2.158	1137	0.598	958.5
1.854	1099	.315	929.8
1.058	1033	.212	916.0
		.104	908.7

Walter, 1889

%	n_D
15°	
0	1.3334
2.10	.3365
8.12	.3449
13.3	.3522
25.6	.3696

Wasastjerna, 1920

M	n H_C	n D	n H_F
18°			
0.0000	1.33168	1.33348	1.33764
.0824	.33355	.33544	.33955
.2120	.33652	.33837	.34264
.3301	.33911	.34101	.34533
.4227	.34110	.34297	.34734
.6279	.34532	.34718	.35171
.6838	.34645	.34827	-
.8449	.34965	.35157	.35613
.9642	-	-	.35848
25°			
0.0000	1.33108	1.33291	1.33701
.0823	.33293	.33478	.33898
.2116	.33581	.33766	.34192
.3295	.33833	.34028	.34463
.4219	.34034	.34222	.34657
.6266	.34455	.34643	.35089
.6823	.34557	.34751	-
.8430	.34876	.35072	.35533
.9620	-	-	.35763

de Garcia, 1920

N	n_D	N	n_D
	17°		
0.06	1.3344	0.5	1.3395
.12	.3349	1	.3442
.25	.3361	2	.3543

Kohlrausch, 1879

%	κ
	18°
5	485
10	932

Mc Gregory, 1894

N	λ	N	λ
	18°		
0.0001	1191	0.01	1072.3
.0002	1168	.05	954.4
.0006	1164.8	.1	898.6
.001	1156.4	.5	763.9
.002	1142.6	1.0	699.5
.006	1098.2	1.5	650

Water (H_2O) + Potassium hemiacid oxalate ($KC_4H_3O_8$)

Benrath, 1942

%	f.t.	%	f.t.
40	116	50	143
45	131	53.2	150

Water (H_2O) + Potassium acid oxalate (KC_2HO_4)

Koppel and Cahn, 1908

%	f.t.	%	f.t.
0.98	-0.25	10.67	60
1.25	0	41.90	103.5 b.t.
4.12	30		

Water + Potassium malonate ($K_2C_3H_2O_4$)

Tammann, 1885

%	p
	100°
12.11	738.0
21.10	709.7

Water + Potassium succinate ($K_2C_4H_4O_4$)

Tammann, 1885

%	p
	100°
13.86	731.0
23.81	698.0
33.10	652.8
39.23	611.7
49.73	521.4
52.02	502.2

Marshall and Cameron, 1907

mol%	f.t.
5.86	0
10.78	20
11.45	25
13.12	40

Wasastjerna, 1920

%	d	
	18°	25°
0	0.99862	0.99707
14.336	1.08977	1.08759
20.832	1.19859	1.19568

de Garcia, 1920

N	d	N	d
	19°		
0.06	1.002982	1	1.051650
.12	.006218	2	.103485
.25	.012854	4	.200310
.50	.025958		

Wasastjerna, 1920

%	n_D H_C	n_D D	n_D H_F
		18°	
0	1.33168	1.33348	1.33764
14.336	.35247	.35449	.35902
29.832	.37635	.37551	.38345
		25°	
0	1.33108	1.33291	1.33701
14.336	.35159	.35362	.35811
29.832	.37551	.37760	.38253

de Garcia, 1920

N	n_D	N	n_D
	19°		
0.06	1.3337	1	1.3460
.12	.3345	2	.3579
.25	.3361	4	.3800
.50	.3393		

Water + Potassium lactate ($KC_3H_5O_3$)

Dietz, Degering and Schopmeyer, 1941

%	b.t./742	f.t.
0	99.33	0
1	-	-
2	-	-
5	99.68	-1.6
10	100.13	-3.7
20	101.35	-8.6
30	103.11	-16.1
40	105.67	-28.1
50	109.35	-51.0
60	114.58	-
70	123.12	-
78.04	-	-

%	d	η	n_D	σ
	25°			29°
0	0.99707	894	1.3329	-
1	1.0018	912	.3338	-
2	.0068	934	.3350	71.4
5	.0217	999	.3390	-
10	.0465	1128	.3456	-
20	.0985	1490	.3589	-
30	.1528	2105	.3729	-
40	.2102	3289	.3870	66.4
50	.2699	5492	.4018	66.4
60	.3307	11320	.4161	65.4
70	.3928	33456	.4306	63.4
78.04	.4426	-	.4417	

Water + Potassium malate ($K_2C_4H_4O_5$)

Schneider, 1881

%	d	$(\alpha)_D$	%	d	$(\alpha)_D$
		20°			
0	0.9982	-	39.133	1.2878	-4.59
9.379	1.0608	-6.81	49.899	.3845	-3.57
22.132	1.1520	-6.06	62.044	.5057	-2.21

Thomsen, 1882

%	$(\alpha)_D$ 10°	$(\alpha)_D$ 20°	$(\alpha)_D$ 30°
16.29	-5.62	-6.35	-7.09
23.25	-5.18	-5.90	-6.57
33.86	-4.48	-5.22	-5.85

Water + Potassium acid malate ($KC_4H_5O_5$)

Schneider, 1881

%	d	$(\alpha)_D$
	20°	
0	0.9982	-
9.427	1.0490	-5.67
13.731	1.0739	-5.29
17.708	1.0964	-5.31
21.683	1.1205	-5.01
26.867	1.1525	-4.70

Water + Potassium tartrate ($K_2C_4H_4O_6$)

Tammann, 1885

%	p	%	p
	100°		
16.35	734.5	50.51	615.4
29.48	704.3	56.61	578.5
39.81	670.0	50.51	615.4

Lowry and Morgan, 1924

t	p	t	p
	sat.sol.		
15.5	9.9	28.8	21.6
17.4	11.5	35.2	30.9
23.4	16.5	40.2	40.3

Legrand, 1835

%	b.t.	%	b.t.
21.18	101	63.78	109
32.06	102	66.23	110
39.32	103	68.43	111
45.14	104	70.41	112
50.02	105	72.18	113
54.23	106	73.79	114
57.86	107	74.75	114.67
61.01	108		

Gerlach, 1886

%	b.t.	%	b.t.
0	100	56.90	108
14.67	101	59.68	109
25.45	102	62.04	110
33.72	103	64.23	111
40.26	104	66.21	112
45.55	105	68.04	113
49.91	106	69.78	114
53.70	107	71.11	115

Fricke and Schützdeller, 1924

m	f.t.	m	f.t.
1.781	-6.569	0.810	-3.141
.381	-5.099	.528	-2.122
.106	-4.190	.227	-1.965

Kremers, 1856

%	d	%	d
	19.5°		
0	0.998	33.66	1.247
4.81	1.030	38.47	.288
9.62	.061	43.28	.333
14.43	.095	48.09	.378
19.23	.131	52.90	.424
24.04	.168	57.71	.474
28.85	.206	62.52	.530

Gerlach, 1869

%	d	%	d
	17.5°		
0	0.9987	40	1.2913
10	1.0636	50	.3797
20	.1335	sat.sol.	.494
30	.2095		

Krecke, 1872

t	d
19.23%	
0	1.1272
25	.1215
50	.1104
75	.0977
100	.0876

Thomsen, 1886

%	d 15°	d 20°	d 25°
0	0.9991	0.9982	0.9971
9.07	1.0605	1.0595	1.0580
18.09	.1270	.1255	.1235
36.39	.2790	.2780	.2770
54.54	.4580	.4500	.4480

Moore, 1895-96

M	d	M	d
	18°		
0.00	0.9987	0.7345	1.1036
.1815	1.0267	1.48	1.2072
.363	1.0525		

Pribram and Glücksmann, 1898

%	d	%	d
	20°		
0	0.99823	18.6567	1.12845
0.6786	1.00231	20.9965	.14637
1.2788	.00678	25.1239	.17851
2.6234	.01575	26.8056	.19212
5.2183	.03294	31.9481	.23417
7.6981	.04995	39.0736	.29686
8.8502	.05773	48.4593	.38367
11.9181	.07945	53.4296	.43452
15.6501	.10622		

Patterson, 1904

t	d	t	d	t	d
5.429%		9.974%		25.647%	
18.16	1.03484	12.44	1.06763	18.97	1.18262
29.89	.03133	24.65	.06384	27.8	.1786
49.05	.0235	50.6	.0528	48.7	.1680
78.5	.0074	81.4	.0354	76.8	.1513
30.89%		39.09%		56.53%	
10	1.2304	17.7	1.2970	20.5	1.4654
28.66	.2214	29.93	.2906	42	.4525
49	.2103	48.45	.2798	79.3	.4290
78	.1926	81	.2605		
62.61%					
27.7	1.5276				
47.8	.5153				
81	.4946				

de Garcia, 1920

N	d	N	d
	18°		
0.06	1.004350	1	1.070728
.12	.008970	2	.137900
.25	.017884	4	.266465
.50	.035892		

Kantele, 1922

N	d 10°	20°	30°	40°	50°
0	0.99973	0.99823	0.99567	0.99224	0.98807
0.12	1.00724	1.00596	1.00338	0.99940	0.99574
0.25	.01515	.01332	.01069	1.00714	1.00292
0.50	.02999	.02834	.02519	.02152	.01666
1	.06012	.05769	.05455	.05043	.04580
1.5	.08835	.08535	.08198	.07771	.07289
2	.11674	.11350	.10946	.10516	.09968
4	.22456	.22018	.21538	.21021	.20390

Fricke and Schützdeller, 1924

M	d	M	d
	at room temperature		
1.535	1.2088	0.759	1.1080
.230	.1693	.506	.0719
.011	.1422	.223	.0316

N	d
25°	
3.328	1.1573
0.931	.0630
.466	.0307
.186	.0107
.093	.039

N	d
25°	
0.093	1.0009
.186	.0077
.466	.0277
.931	.0598
2.328	.1539

Guillaume, 1946

%	d
20°	
13.13	1.0879
40.2	1.2945

Moore, 1895-96

M	η(water=1)	M	η(water=1)
	18°		
0.00	1.000	0.7345	1.342
.1815	.057	1.0	1.489
.25	.080	1.48	1.866
.363	.131	1.5	1.883
.50	.195		

Kantele, 1922

N	η 10°	20°	30°	40°	50°
0	1307	1005	802	655	551
0.12	1330	1020	810	670	560
0.25	1350	1040	830	680	570
0.50	1400	1080	860	700	590
1	1500	1170	930	770	640
1.50	1610	1270	1020	830	700
2	1760	1370	1100	910	760
4	2660	2050	1640	1340	1120

Fricke and Schützdeller, 1924

N	η
25°	
3.325	1343
0.931	1055
.466	969.7
.186	924.4
.093	916.3

de Garcia, 1920

N	n_D	N	n_D
18°			
0.06	1.3341	1	1.3485
.12	.3350	2	.3625
.25	.3372	4	.3913
.50	.3408		

Guillaume, 1946

%	$n_{5780\,\text{Å}}$	(α)magn.10^6
20°		
13.13	1.3524	3.758
40.2	1.3942	3.297

Urban, 1932

t	U 2m	1m	0.5m	0.1m
10	0.7568(12°)	0.8388(11°)	0.9104(14°)	0.9553
15	.7611	.8462	.9116	.9587
20	.7630	.8473	.9139	.9597
25	.7639	.8484	.9147	.9607
30	.7648	.8496	.9155	.9617
35	.7658	.8507	.9164	.9627
40	.7667	.8518	.9172	-

Rotatory power

Britton and Jackson, 1934

M	$(\alpha)_{5461}$	M	$(\alpha)_{5461}$
25°			
0.05	7.00	0.6	7.55
.1	7.05	.7	7.60
.2	7.15	.8	7.65
.3	7.28	.9	7.68
.4	7.37	1.0	7.70
.5	7.46		

Peyches, 1936

M	$(\alpha)_{Hg\ gr}$	M	$(\alpha)_{Hg\ gr}$
20°			
0.001	46.4	0.300	50.0
.004	46.7	.500	50.9
.010	47.0	.750	51.5
.020	47.5	1.000	52.1
.050	48.0	.250	52.5
.100	48.6	.500	52.8
.200	49.4	2.000	53.3

Krecke, 1872

t	(α)mol C	D	E	b	F
19.23%					
0	22.320	27.223	32.925	34.825	40.187
25	22.038	26.836	32.948	34.963	39.920
50	21.514	26.400	32.928	34.693	39.767
75	21.623	26.467	33.030	34.793	40.398
100	21.771	26.415	33.191	34.716	40.412

Thomsen, 1886

%	$(\alpha)_D^{mol}$ 15°	20°	25°
9.07	28.34	28.49	28.65
18.09	29.02	29.19	29.26
36.39	30.07	30.06	30.01
54.54	30.70	30.67	30.57

Pribram and Glücksmann, 1898

%	$(\alpha)_D^{mol}$	%	$(\alpha)_D^{mol}$
20°			
0.6786	27.13	18.6567	29.02
1.2788	27.18	20.9965	29.14
2.6234	27.42	25.1239	29.39
5.2183	27.78	26.8056	29.48
7.6981	28.15	31.9481	29.74
8.8502	28.28	39.0736	29.91
11.9181	28.54	48.4593	30.15
15.6501	28.81	53.4296	30.26

Water + Potassium methyl tartrate ($C_5H_7KO_6$)

Patterson, 1904

%	d	$(\alpha)_D^{mol}$
	20°	
0	0.99825	46.70
5.3067	1.0249	45.99
15.1501	.07625	44.00
24.9555	.13102	42.02
47.085	.26995	37.54
100	-	24.8

t	d	t	d	t	d	t	d
5.3067%		15.150%		24.9555%		47.085%	
16.7	1.0255	16.6	1.0774	13.6	1.1336	15.1	1.2727
30.0	.0220	26.3	.0739	26.7	.1279	31	.2631
47.6	.0146	40.2	.0679	49.3	.1162	48	.2519
63.9	.0062	59.1	.0580	56.1	.1120	63.1	.2411
78.0	0.9981	64.5	.0548	82.3	.0949	79.2	.2297
-	-	77.5	.0468	-	-	-	-

t	$(\alpha)_D$	t	$(\alpha)_D$	t	$(\alpha)_D$	t	$(\alpha)_D$
5.3067%		15.150%		24.9555%		47.085%	
10.6	22.33	7.7	21.02	10.6	20.28	10.1	17.76
14.7	22.50	15.0	21.52	15.1	20.54	17.7	18.32
17.2	22.61	20.7	21.83	35.6	21.52	24.3	18.90
21.5	22.86	27.8	22.08	52.2	21.88	33.0	19.29
30.0	23.19	39.9	22.49	63.2	22.03	70.4	20.05
41.2	23.48	52.9	22.66	99	21.57	99	20.03
52.6	23.47	75.7	22.52				
65.7	23.37	98	22.19				
99	22.55						

Water + Potassium ethyl tartrate ($KC_6H_9O_6$)

Patterson, 1904

%	d	$(\alpha)_D^{mol}$
	20°	
0	0.99825	50
5.06845	1.02117	59.06
15.1035	.06843	56.95
34.747	.16862	53.68
46.787	.2339	51.02
100	-	41.7

t	d	t	d
5.06845%		15.1035%	
13.3	1.0226	13.1	1.0706
22.05	.0206	30.51	.0643
31.37	.0177	55.3	.0518
47.9	.0114	80	.0369
79.2	0.9937		
34.757%		46.787%	
13.3	1.1720	31.7	1.2271
32.42	.1616	52.0	.2130
47.9	.1524	63.5	.2050
64.7	.1413	79.3	.1933
79	.1314		

t	$(\alpha)_D$	t	$(\alpha)_D$
5.06845%		15.1035%	
8.1	26.87	7.6	25.81
14.9	27.14	15	26.19
23.2	27.44	23.6	26.50
40.4	27.73	39.4	26.70
61.1	27.34	56.3	26.68
75.4	26.92	78.3	26.29
98	26.16	98	25.65
34.747%		46.787%	
7	24.23	14.6	23.42
19.7	24.86	25.1	23.75
37	25.31	40.2	24.33
54.6	25.42	56.8	24.49
74.2	25.13	72.2	24.39
99	24.63	98.6	24.06

Water + Potassium propyl tartrate ($KC_7H_{11}O_6$)

Patterson, 1904

t	d	t	d	t	d
4.98903%		25.01%		48.4602%	
18.3	1.01954	13.3	1.1101	13.4	1.2224
22.05	.01856	33.2	.1003	29.8	.2116
29.65	.01623	47.1	.0926	50.3	.1967
49.4	.0082	61.7	.0836	80.3	.1751
79.2	0.9919	81.8	.0697		

t	$(\alpha)_D$	t	$(\alpha)_D$
4.98903 %			
8.9	26.96	36.8	27.44
24.1	27.25	71.4	26.80
50.6	27.23	98.4	25.87
25.01 %			
12.4	25.59	51.9	26.18
21.9	25.93	69.9	25.83
48.4602 %			
7.8	23.89	53.2	25.03
8.2	24.00	63.1	25.05
23.7	24.65	70.2	25.04
38.1	24.86	99.0	24.62

%	d	$(\alpha)_D^{mol}$
	20°	
0	0.99825	64.7
4.98903	1.01912	63.63
25.01	1.10686	60.48
48.4602	1.21842	57.3
100	-	51.0

Patterson, 1904

t	*$(\alpha)_D$
5.6450%	(1+2)= 5.429% anhydride salt
16.1	26.75
29.9	27.07
50.35	27.27
78.7	27.33
99	27.24
10.3707%	(1+2)= 9.979% anhydride salt
12.7	27.06
21.8	27.21
71.3	27.65
98.5	27.60
26.6665%	(1+2)= 25.647% anhydride salt
17.4	28.26
34.2	28.39
47.4	28.47
65.5	28.38
99	28.10
32.12%	(1+2)= 30.89% anhydride salt
18	28.39
37.2	28.58
55.7	28.60
98	28.04
40.645%	(1+2)= 39.09% anhydride salt
17	28.62
29.7	28.69
68	28.50
99	27.76
58.777%	(1+2)= 56.53% anhydride salt
28.2	29.1
16.3	29.41
63.8	28.34
55.3	28.54
99	27.31
65.1%	(1+2)= 62.61% anhydride salt
16.1	69.36
36.4	68.33
99	63.54

*$(\alpha)_D$ refer to the hydrated salt

Water + Sodium-potassium tartrate ($NaKC_4H_4O_6$)

Lowry and Morgan, 1924

t	p dissoc.(4+1)	t	p dissoc.(4+1)
18.0	5.5	29.7	13.4
21.4	7.1	30.1	13.7
24.0	8.6	35.1	19.6
25.1	9.3	40.0	27.9

t	p	t	p
	sat.sol.		
8.3	7.1	35.6	36.7
11.8	9.1	40.3	48.2
12.4	9.5	42.9	49.8
13.8	10.2	43.6	57.2
18.7	14.4	44.9	66.3
20.4	16.1	45.1	59.5
22.5	17.7	46.1	56.4
23.5	19.2	47.6	62.1
24.1	19.8	49.0	61.3
24.6	20.3	52.1	73.9
29.7	26.6	54.3	82.8

N.B. Above 40°, the solution is satured in Rochelle salt with a mixture of sodium and potassium tartrate .

Gerlach, 1886

%	b.t.	%	b.t.
0	100	77.37	118
14.75	101	78.54	119
25.65	102	79.59	120
33.90	103	80.54	121
40.52	104	81.41	122
45.89	105	82.21	123
51.58	106	82.94	124
54.13	107	83.61	125
57.41	108	87.03	130
60.32	109	89.53	135
63.10	110	91.51	140
65.52	111	93.46	145
67.63	112	95.24	150
69.61	113	96.89	155
71.43	114	98.52	160
73.15	115	99	165
74.68	116		
76.07	117		

Raekke, 1913-1914

%	f.t.	%	f.t.
9.09	-1.86	24.24	0
9.12	-1.875	31.51	+9.7
16.11	-3.38	38.60	18.0
20.63	-4.34	50.73	29.5
22.00	-4.70		

%	f.t."rac"	%	f.t. "rac"
26.72	-5.95	36.71	+9.7
27.73	-6.33	47.92	29.5
29.78	-6.98		

Kremers, 1856

%	d	%	d
	19.5°		
0	0.998	22.34	1.160
3.72	1.023	36.06	.191
7.45	.048	29.78	.222
11.17	.076	33.50	.253
14.89	.103	37.23	.285
18.61	.132	40.95	.319

Krecke, 1872

t	d
20%	
0	1.1067
25	.1008
50	.0893
75	.0773
100	.0669

Gerlach, 1886

%	d
17.5°	
0	0.9987
7.448	1.0496
14.896	.1036
22.344	.1605
29.792	.2214
37.240	.2873

Long, 1888

%	d
20°	
0	0.9982
4.88	1.0261
13.97	.0739
22.32	.1202
30.03	.1655
37.18	.2100

Moore, 1895-96

M	d	M	d
	18°		
0.00	0.9987	0.789	1.1087
.20	1.0273	1.656	.2112
.40	.0547		

Grossmann and Wieneke, 1906

t	d	t.	d
	25.017%		
10	1.1344	60	1.1103
20	.1303	70	.1042
30	.1258	80	.0978
40	.1209	90	.0915
50	.1159		

Winther, 1907

%	d 20°	30°	40°	50°	60°
0	0.9982	0.9957	0.9922	0.9881	0.9832
9.35	1.0470	1.0437	1.0398	1.0355	1.0303
24.90	.1313	.1269	.1223	.1173	.1116
35.88	.1965	.1913	.1860	.1805	.1752
42.91	.2401	.2350	.2293	.2240	.2178
46.92	.2417	.2178	.1752	.1116	.0303

Rakshit, 1925

%	d
20°	
1	1.00406
5	.02563
10	.05203
30	.15190
50	.23431

Moore, 1895-96

M	η (water=1)	M	η (water=1)
	18°		
0.00	1.000	0.50	1.252
.20	.092	.789	.476
.25	.112	1.000	.679
.40	.188	.656	2.488

Long, 1888

%	(α)$_D$
20°	
4.88	22.14
13.97	22.16
22.32	22.12
30.03	22.13
37.18	22.06

Krecke, 1872

t	(α) C	D	E	b	F
			20%		
0	18.261	21.820	25.963	27.245	31.569
25	18.524	22.421	26.490	27.667	32.076
50	19.066	22.828	27.355	28.226	32.489
75	19.396	23.513	27.771	28.567	32.977
100	19.762	23.993	28.258	28.980	33.355

Grossmann and Wiencke, 1906

t	(α)$_D$	t	(α)$_D$
	25.017%		
10	21.7	60	22.2
20	22.0	70	22.3
30	22.2	80	22.3
40	22.2	90	22.3
50	22.2		

Winther, 1907

t	(α) red	t	(α) yellow	t	(α) green	t	(α) pale blue	t	(α) dark blue
					9.35%				
15.3	23.24	15.2	29.55	15.8	35.87	15.7	47.83	15.5	52.66
28.9	24.01	28.0	29.80	29.0	36.82	29.6	48.29	30.0	53.43
41.1	24.52	39.6	30.19	42.3	37.27	42.0	49.32	41.5	54.19
53.0	24.50	52.4	30.35	53.6	37.18	53.3	49.56	53.1	54.73
63.2	24.76	63.2	30.51	62.3	37.61	63.0	50.07	62.8	54.99
					24.90%				
16.8	23.76	17.6	29.72	17.3	36.62	17.1	48.28	17.0	52.71
29.0	23.96	30.5	29.90	30.8	36.75	29.5	48.52	30.2	53.06
40.2	24.09	40.5	30.02	42.1	36.87	41.8	48.79	41.1	53.58
50.2	24.22	53.0	30.14	53.0	36.95	52.5	48.97	51.5	54.11
60.9	24.29	61.5	30.17	61.7	37.11	61.3	49.03	61.5	53.99
					35.88%				
16.9	23.59	16.2	29.55	16.3	36.36	17.1	48.28	17.0	52.71
31.7	23.86	28.6	29.76	29.4	36.55	31.0	48.18	30.0	52.80
40.8	23.76	39.9	29.78	40.7	36.54	41.5	48.27	41.8	52.95
51.1	23.81	51.7	29.77	51.1	36.61	51.0	48.34	51.0	53.13
61.5	23.79	63.2	29.74	63.0	36.59	62.3	48.36	62.8	53.15
					42.91%				
17.3	23.46	18.0	29.47	17.9	36.10	17.8	47.59	17.6	52.30
28.3	23.57	31.6	29.56	30.8	36.22	28.9	47.76	30.1	52.47
41.4	23.63	40.0	29.57	38.2	36.25	39.2	47.94	40.5	52.65
52.1	23.66	49.0	29.48	50.7	36.25	51.5	48.05	50.7	52.68
61.0	23.58	59.7	29.49	61.0	36.24	61.0	48.00	61.0	52.70
					46.92%				
16.2	23.21	16.1	29.17	16.5	35.75	16.4	47.23	16.3	51.43
30.1	23.27	28.9	29.17	30.0	35.84	30.8	47.35	31.2	51.63
38.7	23.31	40.3	29.19	40.1	35.81	37.8	47.28	38.7	51.68
51.8	23.34	49.5	29.15	49.9	35.83	50.5	47.35	51.3	51.75
61.4	23.19	61.4	29.07	61.9	36.76	62.0	47.14	62.0	51.75

Water + Potassium citrate ($K_3C_6H_5O_7$)

Tammann, 1885

%	p	%	p
		100°	
16.46	737.5	46.33	623.4
27.70	710.3	58.41	525.8
36.57	676.3	61.06	496.4

Fricke and Schützdeller, 1924

m	f.t.
0.097	-0.528
0.285	-1.475
0.435	-2.220
0.669	-3.470
0.914	-4.950
0.981	-5.364
1.264	-7.264

Wasastjerna, 1920

M	d	M	d
18°		25°	
0.0000	0.99862	0.0000	0.99707
.1491	1.02898	.1488	1.02711
.3466	.06763	.3459	.06550
.4871	.09453	.4861	.09218
.6868	.13173	.6853	.12927
.8208	.15630	.8183	.15359
1.0404	.19569	1.0379	.19284
.3554	.25092	.3522	.24792
.4457	.26633	.4423	.26337

de Garcia, 1920

N	d	N	d
		16.5°	
0.06	1.004106	1	1.065000
.12	.008220	2	.129270
.25	.016621	4	.245170
.50	.033091		

Fricke and Schützdeller, 1924

N	d	N	d
	25°		
0.288	1.0139	2.481	1.1536
0.828	.0501	2.640	.1630
1.243	.0764	3.294	.2018
1.8645	.1157		

N	η	N	η
	25°		
0.288	955.15	2.481	1593
0.828	1091	2.640	1639
1.243	1188	3.294	1958
1.8645	1383		

Wasastjerna, 1920

M	n_{H_C}	n_D	n_{H_F}
		18°	
0.0000	1.33168	1.33348	1.33764
.1491	.33849	.34040	.34463
.3466	.34709	.34908	.35348
.4871	.35294	.35492	.35935
.6868	.36096	.36303	.36765
.8202	.36613	.36820	.37291
1.0404	.37448	.37663	.38151
.3554	.38588	.38810	.39312
.4457	.38911	.39135	.39638
		25°	
0.0000	1.33108	1.33291	1.33701
.1488	.33773	.34463	.34387
.3459	.34625	.35348	.35256
.4861	.35208	.35935	.35850
.6853	.36009	.36765	.36674
.8183	.36522	.37291	.37197
1.0379	.37360	.38151	.38056
.3522	.38501	.39312	.39220
.4423	.38821	.39638	.39541

de Garcia, 1920

N	n_D	N	n_D
	16.5°		
0.06	1.3340	1	1.3478
.12	.3350	2	.3618
.25	.3370	4	.3867
.50	.3410		

Water + Potassium benzoate ($KC_7H_5O_2$)

Tammann, 1885

%	p
100°	
21.18	719.1
31.84	689.3
38.66	665.1
47.87	623.7

Pajetta, 1906

%	f.t.
(3+1)	
41.1	17.5
42.4	25.0
44.0	33.3
46.6	50.0

Sidgwick and Ewbank, 1921

%	f.t.	%	f.t.
5.04	-0.86	44.92	+41.0
9.70	-2.14	50.99	81.0
16.23	-4.29	53.50	97.5
24.23	-7.31	58.42	131.0
39.89	+8.5	66.09	181.0
40.60	13.0		

Guillaume, 1946

%	d	n_{5780}	* $(\alpha)^{5780}_{magn.}10^6$
20°			
20.02	1.0864	1.3761	4.354

* in radians, gauss, centim.

Water + Potassium-o-Hydroxybenzoate ($KC_7H_5O_3$)

Sidgwick and Ewbank, 1921

%	f.t.	%	f.t.
10.49	-1.82	53.33	+24.0 (1+1)
18.82	-3.47	55.82	28.5 "
31.15	-6.52	61.31	61.0
35.80	-8.00	68.97	103.8
44.08	0 (1+1)	70.20	108.5
49.19	+9.0 "	74.80	138.2

Water + Potassium-m-Hydroxybenzoate ($KC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
12.39	-2.41	59.04	+10.0
22.25	-5.06	61.94	33.5
31.34	-8.59	69.60	95.0
49.39	-19.92	75.02	136.8

Water + Potassium-p-Hydroxybenzoate ($KC_7H_6O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
8.04	-1.43	56.70	+64.4 (3+1)
15.55	-3.24	59.34	70.5 "
29.91	+15.8 (3+1)	63.01	86.8
35.50	25.8 "	64.95	129.5
45.71	43.0 "		

Water + Potassium saccharinate ($KC_6H_{11}O_6$)

Rimbach and Heiten, 1908

c	(α) red	yellow	green	pale blue	dark blue
			20°		
10.02	-3.59	-5.54	-7.23	-8.98	-12.27
11.96	-4.05	-5.73	-7.48	-9.24	-12.42
16.00	-4.23	-5.75	-7.50	-9.29	-12.68
20.00	-4.37	-6.10	-8.17	-9.97	-12.85
23.00	-4.52	-6.25	-8.43	-10.15	-13.24
25.05	-4.57	-6.39	-8.64	-10.22	-13.47

t	c	(α) red	yellow	green	pale blue	dark blue
20	25.05	-4.57	-6.39	-8.54	-10.22	-13.47
25	24.99	-3.70	-5.72	-7.90	-9.35	-12.52
30	24.93	-3.45	-5.02	-7.06	-8.53	-11.56
35	24.86	-3.26	-4.83	-6.46	-7.82	-10.36
40	24.79	-2.84	-4.26	-6.07	-7.22	-9.62
45	24.72	-2.45	-3.90	-5.30	-6.51	-9.10
50	24.64	-1.97	-3.25	-4.48	-5.66	-7.75
55	24.56	-1.71	-2.95	-4.19	-5.23	-6.88
60	24.47	-1.35	-2.40	-3.29	-4.58	-6.25

Water + Potassium mellate ($K_6C_{12}O_{12}$)

Guillaume, 1946

%	d^{20}	t	n 5780 Å	*(α)magn. .10⁶)
10.0	1.1009	20	1.3602	3.904
22.17	1.2200	17	1.3778	3.773

* in radians, gauss, centim.

Water + Potassium salt of trinitroxyphenylmethylnitramine ($KC_7H_4O_9N_5$)

Roozeboom, 1889

t	% L_1	L_2	f.t.
10	8.21	72.73	3.13
20	14.34	65.12	-
30	30.17	57.43	5.83
34.5	50	50 C.S.T.	-
50	-	-	22.10
61	-	-	66.12
71	-	-	73.93
74	-	-	72.60

Water + Potassium methyl sulfate (KCH_3SO_4)

Illingworth and Howard, 1884

%	f.t.	%	f.t.
10	-2.3	39.84	-11.8 E
15	-3.6	40	-11.5
20	-5.0	47.08	0.0
30	-8.0	54.8	+12.3

Dixon and Taylor, 1910

mol%	d	n_D
	13.5°	
0	0.9993	1.33346
9.21	1.2914	1.36905

Water + Potassium ethyl sulfate ($KC_2H_5SO_4$)

Illingworth and Howard, 1884

%	f.t.	%	f.t.
10	-2.2	45.01	-14.2 E
20	-4.9	50	-6
30	-8.2	53.71	0
40	-12.1	62.35	+15

Dixon and Taylor, 1910

mol %	t	d	n_D
0	13	0.9994	1.33352
6.45	12	1.2033	1.36632
7.59	14	1.2315	1.36960
7.59	13	1.2315 (sic)	1.37024

Water + Potassium amyl sulfate ($KC_5H_{11}SO_4$)

Illingworth and Howard, 1884

%	f.t.	%	f.t.
10	-1.9	25	-4.8
20	-4.3	33.44	0.0
24.03	-5.4 E	59.46	+17.3

Water + Potassium Dodecylsulfonate ($KC_{12}H_{25}SO_3$)

Vold, 1941

%	tr.t. I	tr.t. II	tr.t. III
50.2	202	35	-
47.2	186	-	177
40.6	112	32	-
22.2	-	31.5	-
15.7	-	31.0	-
10.5	-	30	-
5.58	-	29	-
2.59	-	27	-
1.29	-	24	-
0.65	-	23	-

Water + Potassium-p-Toluene sulfonate ($KC_7H_7O_3S$)

Sinclair, 1933

25°

m	$R.10^5$	m	$R.10^5$
0.1	3274	1.0	2820
.2	3192	.2	2727
.3	3128	.4	2641
.4	3072	.6	2572
.5	3040	.8	2508
.6	3005	2.0	2450
.7	2960	2.5	2316
.8	2907	3.0	2210
.9	2864	3.5	2118

$R = (p_o - p)/mp_o$ where m= molality

Robinson and Stokes, 1949

25°

m	osmotic coefficient
0.1	0.921
.2	.901
.3	.886
.4	.873
.5	.860
.6	.847
.7	.834
.8	.822
.9	.809
1.0	.798
.2	.775
.4	.751
.6	.732
.8	.715
2.0	.700
2.5	.664
3.0	.637
3.5	.615

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.15400	0.1547	0.7125	0.7757
.1722	.1737	.8044	.8851
.2487	.2533	.8754	.9779
.2695	.2752	.9170	1.032
.4099	.4253	1.056	.231
.4241	.4396	.218	.473
.4465	.4637	.401	.757
.4572	.4763	.624	2.126
.4941	.5186	.879	.598
.5732	.6077	.942	.716
.6124	.6536	2.102	3.028
.6240	.6671	.375	.579
.6476	.6971	.516	.870

m_1 = molality of potassium chloride

m_2 = " " " p-toluene-sulfonate

Water + Potassium acid oxalotellurate ($K_2C_2H_6O_{10}Te$)

Rosenheim and Weinheber, 1911

%	f.t.
2.60	0
5.08	20
6.39	30
8.32	40
10.99	50

Water + Potassium brombutyrate ($KC_4H_7O_2Br$)

Timmermans and Van der Haegen, 1933

%	f.t.	%	f.t.	E
4.1	-0.8	37.6	-8.4	-
9.2	-1.6	46.6	-10.8	-
18.1	-4.2	59.2	-13.6	-14.7
24.1	-5.5			

Water+ Potassium borylmalate($KBC_4H_4O_6$)

Jones, 1933

27.29% d^{20}= 1.1390

w.l.	(α)	w.l.	(α)
	20°		
6708	-4.246	5218	-7.116
6439	4.616	5105	7.446
6363	4.745	5086	7.503
6104	5.163	4811	8.427
5893	5.533	4722	8.744
5780	5.766	4602	9.232
5466	6.474	4358	10.35
5461	6.481		

Water + Potassium boryltartrate ($KC_4H_4O_7B$)

Long, 1889

%	d	$(\alpha)_D$
	20°	
0	0.9982	-
4.81	1.0377	10.057
9.27	.0764	8.840
13.44	.1143	8.365
17.32	.1530	8.173

Water + Potassium antimonyltartrate ($KC_4H_4O_7Sb$)

Tammann, 1885

%	p
100°	
10.77	754.8
18.39	750.3
31.02	742.8

XLIX. WATER + SALTS OF OTHER MONOVALENT METALS .

H_2O + Lithium chloride (LiCl)

Heterogeneous equilibria .

Tammann, 1885

t	p 0 %	7.00 %	14.46 %	15.94 %	20.22 %
41.37	59.5	-	51.5	46.6	40.9
43.86	67.8	63.1	58.8	53.0	46.6
46.74	77.4	72.4	67.9	51.5	53.1
49.90	92.1	85.8	80.3	72.0	63.5
51.73	100.8	94.0	87.5	79.0	69.7
54.06	112.9	104.9	98.6	88.5	77.7
56.46	126.6	-	110.7	99.4	87.7
61.66	161.3	150.7	141.0	127.4	112.6
64.95	187.2	175.1	163.6	-	131.1
67.28	207.6	194.1	181.6	164.2	145.6
69.43	228.1	213.6	200.2	181.9	159.5
71.97	254.5	238.7	222.9	202.1	178.1
74.70	286.4	266.8	250.4	226.5	200.3
77.76	324.4	303.3	283.8	253.4	228.0
79.72	351.5	327.8	307.5	278.7	246.8
81.65	379.9	355.7	332.7	302.4	268.3
82.81	397.9	372.4	338.3	316.4	280.7
84.96	433.1	405.4	379.8	344.7	306.3
87.13	471.2	441.0	412.9	375.6	333.3
89.57	517.4	482.9	452.9	412.2	366.5
91.58	558.3	521.4	489.1	446.6	395.5
93.70	604.3	564.5	529.2	-	428.2
96.11	660.4	617.4	578.9	526.8	469.2
97.62	697.6	649.0	610.3	546.7	495.3
99.71	752.1	700.8	658.7	600.1	535.9

%	p	%	p
	100°		
0	760.0	26.09	431.7
4.09	734.3	26.81	418.2
4.17	733.3	30.38	354.9
7.62	705.2	32.19	324.3
7.71	705.0	33.75	293.9
12.78	660.3	34.01	293.4
14.31	629.7	38.98	229.4
20.72	533.4	41.43	191.9
21.98	499.9	49.10	120.2

Dieterici, 1891 (fig.)

%	p	%	p
	0°		
0	4.620	25.14	2.728
7.84	4.280	29.83	2.128
14.53	3.836		

Tower, 1908

%	p	%	p
	0°		
1.12	4.514	18.91	3.371
4.23	4.445	35.06	2.034
8.43	4.251		

Hüttig and Reuscher, 1924

t	p	t	p
(3+1) - (2+1)		(2+1) - (1+1)	
- 35	0.2	0	0.4
--24	0.3	0	0.16-0.4
- 25	0.3	10	1.2
- 16.5	0.8 q	12.5q	2.

t	p	t	p
(1+1)	LiCl	(1+1)	LiCl
30	0.4	65	8.8
35	0.8	70	12.6
40	1.2	70	13.2
45	1.6	70	11.6
30	2.8	70	13.2
55	3.8	70	12.2
60	6.0	100.5 q	90

t	p	t	p
(2+1)	sat. sol.	(2+1)	sat. sol.
- 16.5	0.8	0	0.8
- 15	"	+ 5	"
- 10	"	10	1.0
- 5	"	12.5 q	2.

t	p	t	p
(1+1)	sat.sol.	(1+1)	sat.sol.
12.5	2	45	7.4
15	2.0	50	10 q
20	2.2	50	10.0
25	2.6	55	12.4
30	3.6	58.5	13.8
35	4.6	100.5 q	90
40.5	6.2		

Pearse and Nelson, 1932

m	p	m	p
	25°		
0.00	23.752	4.9582	17.687
0.1004	23.724	5.5749	16.574
0.2185	23.633	6.2074	15.582
0.4040	23.492	6.8567	14.562
0.6082	23.349	7.5199	13.337
0.8142	23.166	8.2002	12.274
0.0216	22.982	9.6109	10.260
1.5471	22.399	11.0955	8.046
2.0845	21.751	12.6577	6.388
2.6327	21.110	14.3221	5.054
3.1940	20.370	16.0508	4.003
3.7687	19.594	17.8635	3.203
4.3576	18.787	19.2186	2.770

Gibson and Adams, 1933

%	p	%	p
	20°28		
43.85	2.313	31.94	6.325
40.49	3.160	26.14	9.257
35.84	4.657	19.32	12.637
35.82	4.704	14.86	14.495

Ueda, 1933

t	p	t	p
	(2+1) + sat. sol. + V		
15	1.75	18	1.96
16	1.83	19	2.07
17	1.90		

t	p	t	p
	(1+1) + sat.sol.		
20	2.21	27.5	3.29
24.5	2.54	30	3.73
25	2.91		

t	p dissoc	t	p dissoc
	(1+1) + LiCl		
20	0.35	30	0.79
25	0.53	35	1.16

Applebey, Crawford and Gordon, 1934

t	p	t	p
	sat.sol.		
60.45	14.9	99.95	80.0
61.35	15.9	109.75	115.6
70.00	26.6	119.85	173.9
70.45	26.2	130.20	242.7
80.15	37.9	140.05	334.4
85.15	46.0	150.15	446.3
90.10	51.2	160.30	615.4
91.00	53.3	170.20	799.0
95.25	62.2	180.20	1035.8
100.30	78.3		

Lannung, 1934

m	p	mol%	p
	18°		
0.725	15.10	0.5905	10.60
0.781	15.08	0.717	9.21
0.823	15.04	0.836	7.93
1.061	14.89	0.960	6.69
1.368	14.70	1.067	5.69
1.649	14.51	1.176	4.87
1.927	14.32	1.334	3.83
2.606	13.80	1.559	2.78
3.394	13.14	1.724	2.24
3.898	12.67	1.168	1.69
4.921	11.65		

t	p dissoc
18	0.27

Johnson and Molstad, 1951

%	p	%	p
	30°		
45.99	3.58	19.35	22.42
44.38	4.39	19.48	24.26
44.39	4.36	13.43	26.35
44.40	4.35	13.50	26.43
44.41	4.33	12.33	27.53
39.82	6.21	12.41	27.35
39.89	6.16	6.54	30.10
36.55	8.21	6.58	30.05
36.63	8.15	3.94	31.18
31.34	12.05	3.95	31.15
31.50	11.97	3.97	31.06
24.93	17.74	3.98	30.95
24.99	17.60	4.00	30.90
25.11	17.55	4.12	30.83
25.37	17.40		
	50°		
48.22	10.37	13.23	78.38
47.47	11.36	13.13	78.56
47.41	11.42	13.35	78.03
43.66	15.06	8.35	85.13
43.58	15.19	8.29	85.14
39.81	20.36	4.99	88.74
39.75	20.47	4.95	88.76
36.95	25.18	3.26	90.22
36.84	25.35	3.23	90.27
30.43	39.29	3.21	90.29
30.29	39.60	3.18	90.32
22.20	59.95	3.16	90.33
22.14	60.28	3.10	90.38
	70°		
45.61	37.44	28.48	114.58
45.71	37.04	18.00	176.33
36.96	68.03	18.13	175.62
37.16	67.39	18.30	175.16
28.35	115.62	18.43	174.61

Gokcen, 1951.

t	p	t	p
	sat.sol.		
23.90	2.63	39.90	7.26
29.90	2.93	44.90	9.82
34.90	5.32	54.84	16.70

Gerlach, 1886

%	b.t.	%	b.t.
0	100	39.94	135
3.38	101	40.48	136
6.54	102	41.10	137
9.09	103	41.69	138
10.11	104	42.28	139
13.04	105	42.86	140
14.89	106	43.50	141
16.67	107	44.13	142
18.03	108	44.75	143
19.35	109	45.35	144
20.63	110	45.95	145
21.88	111	46.67	146
23.08	112	47.34	147
24.24	113	48.06	148
25.09	114	48.72	149
25.93	115	49.37	150
26.74	116	50.00	151
27.54	117	50.60	152
28.31	118	51.22	153
29.08	119	51.80	154
29.83	120	52.49	155
30.55	121	53.15	156
31.27	122	53.76	157
31.97	123	54.06	158
32.69	124	54.45	159
33.33	125	55.05	160
33.99	126	55.66	161
34.64	127	56.23	162
35.27	128	56.79	163
35.90	129	57.45	164
36.51	130	58.06	165
37.11	131	58.76	166
37.70	132	59.42	167
38.27	133	60.06	168
39.32	134		

Johnston, 1906

%	b.t.	%	b.t.
2.64	100.587	27.21	114.013
6.46	101.592	29.15	116.919
9.48	102.547	30.74	118.760
11.65	103.438	32.23	121.860
14.24	104.649	33.93	123.956
16.51	106.017	36.29	128.240
18.87	107.542	42.95	138.272
21.17	109.294	43.87	141.564
22.65	110.419	45.43	144.168
23.59	110.523		
1.01	100.231	14.25	104.649
3.74	100.743	16.51	106.017
6.46	101.592	18.88	107.542
9.38	102.547	21.18	109.294
11.66	103.438	23.60	111.419

Kremers, 1856 and 1857.

%	f.t.	%	f.t.
38.91	0	56.49	95
44.64	20	58.13	140
51.02	65	59.17	160

b.t. (sat.sol.) = 171°

Johnston, 1906

%	f.t.	%	f.t.
34.83	- 73.0	9.14	- 7.90
30.32	- 62.0	4.98	- 3.10
22.54	- 45.5	2.58	- 1.50
12.91	- 12.95		

Rodebush, 1918

%	f.t.	%	f.t.
5.19	- 5.11	13.07	- 18.75
10.68	-12.22	15.70	- 25.44

Hüttig and Reuscher, 1924

%	f.t.	%	f.t.
38.9	0	51.0	65
41.9	10	53.5	80
44.7	20	56.3	96
46.0	30	56.5	100.5 - q
47.4	40.5	58.2	140
48.7	50	59.2	160
			(2+1) (1+1)

Newton, Friend and Colley, 1931

%	f.t.	%	f.t.
40.85	0 (2+1)	45.47	21.2
42.38	8.0	45.89	24.6
42.75	10.2	46.32	29.0
43.32	13.0	44.67	34.5
43.50	13.8	47.47	41.0
43.86	14.2	48.23	47.0
(47.91)	(16.0)	50.05	61.6
44.60	16.2	51.71	72.0
44.23	16.4	53.37	81.6
45.31	18.0 (1+1)	54.60	86.6
45.22	19.0	54.54	88.0

Bassett and Sanderson, 1932

%	f.t.		%	f.t.
40.87	0	(2+1)	47.47	40.
45.85	25	(1+1)	52.71	80.

Friend, Hale and Ryder, 1937

%	f.t.	%	f.t.	
51.6 (1+1)	71.5	57.55	98.2	
51.71	71.0	56.40	101.8	
53.37	81.6	56.52	107.6	0 aq
54.46	87.0	56.95	115.4	
55.27	88.0	57.00	117.0	
55.84	94.6	57.28	123.0	
56.5	95.0	57.58	130.5	
56.0	95.3	58.1	140.0	
56.49	95.5	57.90	140.5	
56.57	96.2	58.46	154.0	
56.60	97.1	59.2	160	

Voskresenskaja and Yanatieva, 1937

wt%	mol%	f.t.	
4.0	1.74	-2.4	
8.0	3.56	-9.0	
14.0	6.46	-23.0	
18.0	8.52	-36.0	
21.0	10.14	-50.0	
24.0	11.84	-62.0	
24.4	12.05	-66.0	E
26.4	13.21	-63.0	
28.2	14.29	-60.4	(5+1)
29.6	15.15	-58.0	
30.4	15.65	-57.0	tr.t.
30.5	15.70	-54.0	(3+1)
30.8	15.89	-48.0	
33.4	17.55	-31.0	
36.4	19.55	-19.2	
37.2	20.09	-15.6	tr.t.
38.8	21.23	0	(2+1)
38.9	21.28	0	
40.0	22.05	4.8	
43.0	24.25	14.0	
43.86	24.90	14.2	
45.6	26.24	20.5	tr.t.
46.0	26.55	25.0	(1+1)
45.8	26.38	25.0	
46.2	26.71	30.0	
46.1	26.63	30.0	

Johnson and Molstad, 1951

%	f.t.
46.22	30
48.75	50
51.10	70

Maran, 1956 (fig.)

%	f.t.	trans.
0	0	-
10	-10	-
20	-45	-
25.0	-75 E	-
32.0	-65.6	(3+2) - (5+2)
35.0	-30	-
38	-20.5	(2+1) - (3+2)
40	0	-
46	+19.0	(1+1) - (2+1)
50	60	-
55	90	-
57	94	B - (1+1)

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	.091	1.6	.939
.2	.116	1.8	.945
.3	.142	2.0	.954
.4	.212	2.5	.963
.5	.286	3.0	.973
.6	.366	3.5	.984
.7	.449	4.0	.995
.8	.533	4.5	1.006
.9	.619	5.0	.018
1.0	.705	5.5	.041
.2	.791	6.0	.066
.4	.939		

Robinson and Stokes, 1949

m	activity coefficient	m	activity coefficient
	25°		
0.1	0.790	1.6	0.853
0.2	0.757	1.8	0.885
0.3	0.744	2.0	0.921
0.4	0.740	2.5	1.026
0.5	0.739	3.0	1.156
0.6	0.743	3.5	1.317
0.7	0.748	4.0	1.510
0.8	0.755	4.5	1.741
0.9	0.764	5.0	2.02
1.0	0.774	5.5	2.34
1.2	0.796	6.0	2.72
1.4	0.823		

Properties of phases .

Density .

Kremers, 1856 and 1857

%	d	%	d
	19.5°		
4.798	1.0261	21.316	1.1229
9.404	1.0523	30.248	1.1812
15.533	1.0878	37.601	1.2341

t	d		
	4.21 %	8.26 %	11.81 %
0	0.99739	0.99671	0.99630
10	0.99828	0.99805	0.99793
19.5	1.00000	1.00000	1.00000
30	1.00278	1.00284	1.00280
40	1.00614	1.00611	1.00599
50	1.01018	1.00991	1.00964
60	1.01479	1.01423	1.01366
70	1.01994	1.01901	1.01811
80	1.02560	1.02418	1.02291
90	1.03178	1.02983	1.02811
100	1.03856	1.03596	1.03375
	16.53 %	19.35 %	35.82 %
0	0.99605	0.99588	-
10	0.99784	0.99783	0.99746
19.5	1.00000	1.00000	1.00000
30	1.00279	1.00278	-
40	1.00584	1.00575	1.00585
50	1.00924	1.00903	-
60	1.01295	1.01259	1.01197
70	1.01701	1.01643	-
80	1.02136	1.02053	1.01863
90	1.02604	1.02494	-
100	1.03098	1.02968	1.02583

Gerlach, 1859 and 1889

%	d	%	d
	15°		
0	0.9991	21	1.123
1	1.005	22	1.130
2	1.011	23	1.135
3	1.017	24	1.141
4	1.023	25	1.147
5	1:029	26	1.154
6	1.034	27	1.160
7	1.039	28	1.167
8	1.045	29	1.174
9	1.050	30	1.1809
10	1.058	31	1.188
11	1.063	32	1.195
12	1.070	33	1.202
13	1.075	34	1.209
14	1.080	35	1.217
15	1.085	36	1.224
16	1.092	37	1.231
17	1.098	38	1.239
18	1.103	39	[illegible]
19	1.109	40	[illegible]
20	1.1162		

Fouqué, 1867

%	$d_{0°}$	d_t
7.63	1.0491	1.0477 (11°)
27.73	1.2495	1.2457 (12°)

Kohlrausch, 1879

%	d	%	d
	18°		
2.5	1.0132	20	1.115
5	1.0274	30	1.181
10	1.0563	40	1.255

Kuschel, 1881

%	t	d
25.46	17.7	1.1524
12.74	20.7	.0735
7.30	18.8	.0418
3.38	17.0	.0179
	16.1	.0195
1.00	15.2	.0060
0.50	14.4	.0032
.43	12.8	.0030
.20	16.2	.0011
.18	17.5	.0008

Bender, 1887

N	d 15°	$\tau \cdot 10^6$ 15-20°	20-25°
0	0.9991	-	-
1	1.0235	1017	1226
2	.0464	1107	1263
3	.0684	1143	1272
4	.0902	1157	1269
5	.1110	1172	1269
6	.1330	1193	1267

Perkin, 1894

mol%	d 15°	25°
23.76	1.3055	1.3056 (20°)
11.89	1.1565	1.1534
7.84	1.10552	1.1028
0	0.9991	0.9971

Lemoine, 1897

%	d	%	d
		0°	
4.26	1.026	32.5	1.203
12.18	1.073	41.4	1.267
22.2	1.133	43.2	1.282

Linebarger, 1899

%	d	%	d
		25°	
0	0.99707	26.14	1.15468
9.18	1.05023	40.49	1.25904
18.06	1.10298		

Forchheimer, 1900

%	d	%	d
		20°	
35.97	1.2230	11.83	1.0691
26.90	1.1550	4.37	1.0232
21.10	1.1215	0	0.9982

Hosking, 1904

t	d 0%	d 0.00645 N	d 0.1030 N
0	0.9999	1.0003	1.0026
10	9997	0.9999	1.0023
15	9992	9992	1.0016
18	9986	9986	1.0012
20	9983	9983	1.0009
30	9958	9958	0.9980
40	9923	9923	0.9947
50	9882	9882	0.9905
60	9834	9834	0.9855
70	9779	9779	0.9803
80	9719	9719	0.9747
90	9656	9656	0.9683
100	9586	9586	0.9614

t	d 0.5203 N	d 1.0125 N	d 2.937 N
0	1.0131	1.0253	1.0694
10	1.0124	1.0243	1.0676
15	1.0115	1.0234	1.0665
18	1.0110	1.0229	1.0658
20	1.0106	1.0225	1.0654
30	1.0083	1.0197	1.0623
40	1.0045	1.0162	1.0590
50	1.0003	1.0121	1.0552
60	0.9958	1.0076	1.0509
70	0.9905	1.0028	1.0464
80	0.9849	0.9972	1.0416
90	0.9790	0.9915	1.0366
100	0.9728	0.9856	1.0312

t	d 5.02 N	d 7.36 N	d 10.71 N
0	1.1142	1.1654	1.2368
10	1.1119	1.1626	1.2334
15	1.1107	1.1611	1.2319
18	1.1099	1.1601	1.2309
20	1.1094	1.1596	1.2301
30	1.1064	1.1565	1.2267
40	1.1032	1.1533	1.2232
50	1.0996	1.1500	1.2196
60	1.0957	1.1462	1.2159
70	1.0918	1.1422	1.2120
80	1.0876	1.1381	1.2080
90	1.0830	1.1340	1.2039
100	1.0781	1.1298	1.1998

Chéneveau, 1907

%	d	%	d
		18°	
32.16	1.1968	9.17	1.0523
25.15	1.1489	0	0.9986
17.52	1.1004		
		19°	
43.44	1.2821	8.39	1.0441
30.31	1.1841	6.77	1.0357
16.12	1.0862	3.44	1.0180
13.10	1.0696	0	0.9984
9.99	1.0524		
		24°	
32.16	1.1959	0	0.9973
		30°	
32.16	1.1947	0	0.9957

Green, 1908

N	d	N	d
		18°	
0	0.9986	4.870	1.1078
0.466	1.0098	6.198	1.1352
0.937	1.0211	8.034	1.1743
2.043	1.0464	9.973	1.2154
3.439	1.0768	12.227	1.2633
		25°	
0	0.9971	5.614	1.12229
0.6159	1.01208	8.698	1.18640
0.9585	1.02000	12.04	1.25680
3.273	1.07173		

Tower, 1908

%	d	%	d
	0°		
4.579	0.9999	4.251	1.0490
4.514	1.0132	3.371	1.1151
4.445	1.0254	2.034	1.2247

Guerdjikova, 1910

%	d	%	d
	25°		
		36.356	1.2285
10.677	1.0623	37.655	-
21.102	.1373	40.755	.2723
29.443	.1818	43.328	.2874
31.504	.2044		

Baxter,Boylston and al., 1911

%	d	%	d
	25°		
0.	0.99707	2.7648	1.01308
0.6254	1.00068	2.7703	1.01311
0.6977	1.00116	2.7352	1.02966
0.9329	1.00240	3.5965	1.03484
0.9578	1.00243	10.5860	1.05832
0.9688	1.00271	13.1099	1.07299
1.4114	1.00511		

Lübben, 1913

N	d	N	d
	18°		
0	0.99862	2.009	1.04576
0.4988	1.01083	3.956	1.08830
1.013	1.02300		

Sakhanov, 1913

%	d	%	d
	25°		
0	0.9971	35.74	1.2230
7.331	1.0389	45.86	1.2983
21.28	1.1251		

Grufki, 1913

N	d
18°	
0	0.99862
0.4971	1.01077
1.000	.02259
1.980	.04509
3.966	.08864

Tucker, 1915

mol%	d	mol%	d
	17°		
22.74	1.2604	8.26	1.0988
18.53	1.2156	6.00	1.0716
13.42	1.1592	4.35	1.0540
10.81	1.1280	0.00	0.9988

Baxter and Wallace, 1916

%	d	%	d
100.0°		70.19°	
38.92	1.21495	25.73	1.13341
21.70	1.0995	11.20	1.04366
11.49	1.03026	5.74	1.01151
4.80	0.99098	2.91	0.99491
1.18	0.96648	1.41	0.98617

%	d	%	d
50.04°		25.0°	
45.53	1.28433	45.29	1.26918
25.59	1.14014	25.45	1.14744
11.12	1.05211	11.04	1.06042
5.69	1.02095	5.65	1.02961
2.88	1.00474	2.86	1.01361
1.40	0.99622	1.99	1.00516

%	d	%	d
	0.0°		
25.33	1.15345	2.85	1.01726
10.99	1.06549	1.39	1.00840
5.62	1.03386		

Alfimoff, 1917

t	d	t	d
c= 60.40		c= 33.96	
34.9	1.2914	0	1.1810
56.3	.2837	21.1	.1758
76.8	.2759	34.3	.1720
		53.6	.1658
			.1576

t	d	t	d
c=20.62		c=6.92	
0	1.1056	0	1.0394
18.3	.1013	19.7	.0364
34.4	.0969	24.8	.0356
54.6	.0902	33.2	.0332
		55.1	.0253
		77.7	.0148

de Block, 1925

%	d	%	d
	16°		
0	0.9990	25.1	1.1475
4	1.0230	34.5	1.2130
14.7	1.0840	41.7	1.2186

Hüttig and Keller, 1925

N	d	N	d
	20°		
11.937	1.25462	1.53468	1.03456
9.6059	1.20597	0.95783	1.02122
6.01017	1.13140	0.602855	1.01307
3.81769	1.08442	0.244046	1.00425
2.45603	1.05508	0.097332	1.00076
1.53722	1.03464		

Fontell, 1927

%	d 6°	25°	50°	70°
0.00	0.99997	0.99707	0.98807	0.97781
2.77	1.01654	1.01300	1.00403	.99384
5.10	.03019	.02623	.01742	1.00749
15.18	.08889	.08439	.07666	.06731
25.30	.15143	.14647	.13826	.13043
35.69	.22446	.21894	.21016	.20233
43.87	.28764	.28096	.27432	.26312

Schreiner, 1928

N	d	N	d
	18°		
0	-	3.014	1.06801
0.509	1.01109	4.840	.10643
1.024	.02327	8.086	.17533
2.012	.04587	9.642	.20754

Kohner, 1928

m	d	m	d
	25°		
0	0.99707	3.00131	1.06211
0.099767	0.99951	4.00274	1.08050
0.49558	1.00836	5.00400	1.05843
1.00109	1.02057	8.96790	1.16166
2.00425	1.04211		

Gibson and Adams, 1933

%	d	%	d
	20.28°		
43.85	1.2822	31.94	1.1908
40.49	1.2538	26.14	1.1520
35.84	1.2186 (sic)	19.32	1.1105
35.82	1.2188	14.86	1.0825

Applebey, Crawford and Gordon, 1934

%	t	d (sat.sol.)
40.90	1.05	1.2678
42.71	10.05	1.2789
44.06	15.70	1.2875
44.86	18.25	1.2926
45.47	24.15	1.2962
47.09	40.05	1.3034
48.23	50.20	1.3080
50.26	64.95	1.3173
52.94	80.85	1.3312
54.82	89.55	1.3418
55.83	95.15	1.3471
56.3	100.2	1.347
56.9	120.2	1.344
57.6	140.3	1.339
58.3	155.6	1.338

Scott, Obenhaus and Wilson, 1934

%	d	%	d
	35%		
42.031	1.26314	18.969	1.10418
34.898	.20970	16.279	.08776
26.021	.14836	9.655	.04943
23.010	.12896	4.1221	.01772

Gibson, 1935

%	d	%	d
	25°		
0.00	0.9970	19.66	1.1109
4.44	1.0227	27.91	1.1630
7.66	1.0422	34.62	1.2098
9.82	1.0524	39.30	1.2445
15.06	1.0838		

Scott and Bridger, 1936

%	d
35°	
41.020	1.25597
37.003	1.22536
27.450	1.15772

Guillaume, 1946

%	d
20°	
6.60	1.0343
31.9	1.1953

Rodnyanskii and Galinker, 1955

t	d 1N	2N	3N
25	1.021	1.042	1.062
50	1.009	1.032	1.054
100	0.984	1.007	1.031
150	0.948	0.974	0.997
200	0.906	0.938	0.963
250	0.861	0.897	0.923
300	0.803	0.847	0.876
340	0.735	0.784	0.815

Pohle, 1906 and 1908

%	π relative	%	π relative
	13.9°		
0	1.000.	11.21	0.785.
3.60	0.980.	15.88	0.710.
7.08	0.857.	19.59	0.746.

Scott, Obenhaus and Wilson, 1934

%	π	%	π
	35°		
42.031	20.12	18.969	29.06
34.898	22.61	16.279	30.52
26.021	26.14	9.655	34.45
23.010	27.41	4.1221	38.35

Gibson, 1935

%	π(1-1000 bars)	%	π(1-1000 bars)
	25°		
0.00	39.35	19.66	27.75
4.44	36.11	27.91	24.61
7.66	34.00	34.62	22.22
9.82	32.64	39.30	20.61
15.06	29.86		

Scott and Bridger, 1936

%	π
35°	
41.020	20.44
37.003	21.86
27.450	25.62

Tammann and Schwarzkopf, 1928

t	Dv.10^4	t	Dv.10^4
	Vo°=1		
19.35%		21.91%	
- 1	- 1.26	-33°3	- 4.21
- 2	2.60	-40°6	7.92
- 4	4.83	-10	12.80
- 7	8.24	-13	15.90
-10	10.85	-16	18.60
-13	13.15	-20	21.75
-17	15.53	-24	24.62
-20	16.30	-28	26.77
-23	17.90	-32	28.28
-24	18.16	-35	29.01
-25	18.36	-39	29.53
-26	18.51	-40	29.59
-28	18.61	-41	29.63
-29	18.56	-42	29.56
-32	18.10	-43	29.54
-34	17.64	-46	29.44
-36	16.84	-50	28.73
-38	15.93	-53	27.91
-40	14.79	-58	26.16
-43	12.72	-62	24.26
	$V_{-20^\circ}=1$		
18.50%		18.75%	
-21	- 0.36	-23	- 1.06
-24	- 0.89	-25	- 1.53
-27	- 1.01	-28	- 2.00
-28	- 0.94	-30	- 1.46
-30	- 0.63	-33	- 1.06
-32	- 0.42	-36	+ 0.42
-34	+ 0.53	-40	+ 2.35
-36	+ 1.38	-42	+ 4.03
-38	+ 2.39	-47	+ 8.70
-40	+ 3.65	-49	+10.84
-42	+ 4.90	-50	+11.86
23.69%		24.24%	
-27	- 4.73	-48.5	-18.09
-29	5.97	-49	18.37
-46	15.51	-50	18.57
-50	16.15	-52	19.18
-54	16.54	-54	19.69
-55	16.56	-57	20.49
-57	16.99	-60	20.88
-62	16.33	-62	21:03
-64	16.15	-64	21.26
-68	15.64	-64.8	21.46
-72	14.16		
	28.87 %		
-30	- 7.64 cc	-43	16.35
-35	11.73	-44	17.04
-39	13.62	-49	20.62
-40	14.29	-50	21.18

Viscosity and surface tension .

Hosking, 1904

t	0%	η 0.00645 N	0.1030 N
0	1794	1802	1806
10	1309	1310	1320
15	1143	1144	1152
18	1060	1060	1070
20	1009	1012	1020
30	802	804	810
40	659	659	668
50	554	554	561
60	472	472	479
70	408	408	414
80	358	358	363
90	319	319	323
100	287	287	292

t	0.5203 N	η 1.0125 N	2.937 N
0	1894	2025	2652
10	1397	1497	1965
15	1220	1308	1718
18	1133	1212	1600
20	1075	1154	1527
30	859	920	1228
40	706	758	1012
50	595	636	853
60	510	542	728
70	441	474	632
80	390	417	558
90	346	368	496
100	308	325	444

t	5.02 N	η 7.36 N	10.71 N
0	3558	5576	12900
10	2652	4150	9360
15	2331	3640	8050
18	2169	3365	7390
20	2070	3204	6990
30	1664	2562	5400
40	1366	2112	4310
50	1156	1751	3552
60	995	1482	2957
70	865	1289	2508
80	764	1133	2161
90	686	1003	1885
100	617	895	1660

Green, 1908

N	η	N	η
	17.82°		
0	1063.0	7.154	3169.5
1.414	1287.2	7.944	3702.7
2.783	1557.0	8.815	4545.0
3.915	1834.0	9.973	6030.0
4.922	2136.5	11.084	8218.9
5.654	2396.4	12.036	10851
6.183	2700.5	12.640	13074
	25°		
0	895.5	6.218	2244.4
0.6175	972.4	6.976	2583.8
1.163	1046.6	8.445	3482.9
2.001	1182.5	8.804	3778.6
2.860	1331.2	9.173	4122.9
4.029	1575.8	10.400	5625
4.970	1818.8	11.265	7144
5.325	1931.9	12.345	9589

Sakhanov, 1913

%	η	%	η
	25°		
0	895	35.74	5664
7.331	1140	45.86	16730
21.28	1930		

Alfimoff, 1917

t	η
60.40 g/100 c	
34.9	11400
56.3	6700
76.8	4520

t	η	t	η
	6.92 g/100 c		
0	2090	33.8	900
19.7	-	55.1	620
24.8	1090	77.7	450

t	η	t	η
c = 33.96		c = 20.62	
0	6040	0	3110
21.1	3430	18.3	1920
34.3	2580	34.4	1350
53.6	1810	54.6	960
76.4	1300	77.2	700

Wolkova and Titow, 1933

d	η	d	η
	25°		
1.2838	13032	1.2049	4514
1.2659	10060	1.1970	4125
1.2632	9766	1.1771	3304
1.2498	7923	1.1630	2878
1.2403	7067	1.1437	2433
1.2268	5956	1.1215	2043
1.2169	5239	1.1043	1777

Jacopetti, 1942

t	η 0.5m	1m	2m	3m	4m
18	1139	1218	1382	1561	1764
25	955	1024	1167	1322	1496
40	696	749	858	976	1107
60	499	540	622	711	808
80	379	410	476	546	620
100	301	325	377	433	496

t	η 5m	6m	7m
18	1999	2274	2607
25	1696	1928	2201
40	1256	1427	1627
60	927	1041	1185
80	704	798	906
100	565	642	727

Linebarger, 1899

%	σ
25°	
0	71.78
9.18	73.17
18.06	78.17
26.14	83.39
40.49	93.88

de Block, 1925

%	σ	%	σ
	16°		
0	73.11	25.1	85.40
4.0	74.65	34.5	91.97
14.7	79.64	41.7	95.55

Harkins and Wampler, 1928

%	Dσ	%	Dσ
0.5	0.81	4.0	6.70
1.0	1.63	8.0	13.53
1.5	2.43	13.0	20.77
2.0	3.26	19.4	28.01

Kremers 1856 and 1857

%	n_D
17°	
0	1.3332
24.53	1.3844
40.08	1.4212

Walter 1889

%	n_D	%	n_D
	15°		
4.20	1.3426	20.3	1.3762
8.76	1.3517	33.0	1.4049

Chéneveau, 1907

%	n_D	%	n_D
18°		19°	
32.16	1.4035	43.44	1.4305
25.15	.3867	30.31	.3990
17.52	.3696	16.12	.3644
9.17	.3519	13.10	.3585
0	.3331	9.99	.3523
		8.39	.3493
		6.77	.3462
		3.44	.3397
		0	.3331
24°		30°	
32.16	1.4029	32.16	1.4023
0	1.3326	0	1.3319

Guerdjikova, 1910

%	n_D	%	n_D
	25°		
0	1.33255	31.504	1.4069
10.677	1.3565	36.356	1.4149
21.102	1.3838	40.755	1.4289
29.443	1.3993	43.328	1.4343

Baxter, Boylston and al., 1911

%	n_D	%	n_D
	25°		
0	1.33246	2.7648	1.33833
.6254	.33388	2.7703	.33842
.6977	.33405	2.7352	.34465
.9329	.33436	3.5965	.34653
.9578	.33446	10.5860	.35449
.9688	.33460	13.1099	.36024
1.4114	.33555		

Müller and Guerdjikoff, 1912

%	n_D	%	n_D
	25°		
0	1.3326	31.50	1.4069
10.68	1.3565	36.36	1.4149
15.27	1.3673	40.76	1.4289
21.10	1.3838	43.33	1.4343

Heydweiler, 1913

N	n_D	N	n_D
	18°		
0.5	1.33774	2.0	1.35053
1.0	1.34207	4.0	1.36651

Grufki, 1913

N	n Hα	n Hβ	n Hγ
0.4971	1.33580	1.34192	1.34525
1.000	1.34011	1.34639	1.34979
1.980	1.34829	1.35490	1.35847
3.966	1.36398	1.37117	1.37512
0	1.33140	1.33735	1.34056

Lübben, 1913

N	n 2144.5 Å	2194.7	2265.1
	18°		
0	1.40414	1.39900	1.39274
0.4988	1.41203	1.40652	1.39984
1.013	1.41995	1.41400	1.40689
2.009	1.43468	1.42800	1.42012
3.956	1.46243	1.45438	1.43505

N	n 2288.1 Å	2312.9
0	1.39089	1.38895
0.4988	1.39786	1.39584
1.013	1.40476	1.40254
2.009	1.41780	1.41538
3.956	1.44229	1.43944

N	n 2573.2 Å	2748.7
0	1.37361	1.36654
0.4988	1.37967	1.37255
1.013	1.38555	1.37782
2.009	1.39685	1.38849
3.956	1.41799	1.40844

Alfimoff, 1917

t	n_D	t	n_D
	60.40gr*		
18.3	1.4373	28.0	1.4357
20.1	.4371	46.0	1.4334
	33.96gr		
18.0	1.3979	35.0	.3960
20.2	.3977	49.0	.3943
29.0	.3966	59.0	.3930
	20.62gr		6.92 gr
17.0	1.3711	17.0	1.3477
20.0	.3708	19.3	.3475
34.0	.3693	20.1	.3474
		40.8	.3456
		53.0	.3443

* concentration in g/cc

Hüttig and Keller, 1925

N	n^{20} 4340A°	4360	4860
11.937	1.43503	1.43479	1.42979
9.6059	.41845	.41818	.41348
6.01017	.39136	.39113	.38698
3.81769	.37347	.37323	.36947
2.45603	.36219	.36196	.35846
1.53722	.35434	.35408	.35077
1.53468	.34923	.34901	.34576

N	n 5460 Å	D	6560
11.937	1.42554	1.42328	1.42052
9.6059	40945	40728	40468
6.01017	38339	38135	37911
3.81769	36221	36444	36225
2.45603	35547	35374	35164
1.53722	34782	34623	34427
1.53468	34291	34138	33944

Fontell, 1927

%	n_{H_α} 6°	25°	50°	70°
0.00	1.33230	1.33115	1.32758	1.32374
2.77	.33816	.33684	.33343	.32992
5.10	.34335	.34182	.33849	.33479
15.18	.36440	.36284	.35984	.35687
25.30	.38679	.38506	.38227	.37968
35.69	.41193	.40954	.40665	.40387
43.87	.43102	.42898	.42579	.42327

%	n_D 6°	25°	50°	70°
0.00	1.33426	1.33298	1.32938	1.32569
2.77	.34011	.33867	.33528	.33127
5.10	.34517	.34361	.34019	.33685
15.18	.36659	.36488	.36221	.35892
25.30	.38926	.38755	:38470	.38185
35.69	.41428	.41210	.40944	.40677
43.87	.43387	.43178	.42856	.42586

%	n_{H_β} 6°	25°	50°	70°
0.00	1.33855	1.33708	1.33358	1.32963
2.77	.34463	.34315	.33975	.33618
5.10	.34985	.34816	.34470	.34119
15.18	.37205	.37007	.36709	.36415
25.30	.39498	.39322	.39036	.38762
35.69	.42085	.41862	.41542	.41272
43.87	.44073	.43835	.43528	.43262

Kohner, 1928

m	n_D	m	n_D
	25°		
0	1.33253	3.00131	1.35646
0.099767	1.33344	4.00274	1.36335
0.49558	1.33694	5.00400	1.36979
1.00103	1.34122	8.96790	1.39241
2.00425	1.34913		

Schreiner, 1928

N	n C	n D	n F
	18°		
0	1.33125	1.33311	1.33725
0.509	.33578	.33766	.34196
1.024	.34022	.34215	.34658
2.012	.34842	.35054	.35508
3.014	.35645	.35856	.36342
4.840	.37036	.37263	.37790
8.086	.38946	.39711	.40297
9.642	.40529	.40792	.41412

Okazaki, 1933

%	Verdet's constant (3441 Å)	%	Verdet's constant (3441 Å)
	28°		
5.67	0.04859	24.13	.06561
12.45	.05511	31.29	.07305
19.98	.06214	33.46	.07477
20.18	.06244	41.88	.08179

%	Verdet's constant (D)	%	Verdet's constant (D)
	25°		
6.629	.01458	27.65	.01948
11.837	.01574	31.43	.02039
15.03	.01643	37.11	.02165
17.78	.01714	40.53	.02238
22.82	.01839		

Rao, 1938

Raman spectra .

Perkin , 1894

mol%	t	$(\alpha)_{magn.}$
23.76	19	1.7294
11.89	23.2	1.4170
7.84	17.4	1.2877

Guerdjikova, 1910

%	(α) magn	%	(α) magn
	25°		
0	5.068	36.356	8.014
10.677	5.778	37.655	-
21.102	6.841	40.755	8.457
29.443	7.447	43.328	8.635
31.504	7.614		

Guillaume, 1946

%	t	n	* $(\alpha)_o$ magn.10^6 5780Å
6.60	21	1.3466	4.240
31.9	22.5	1.4034	5.209

* in radians ,gauss, centim.

Forchheimer, 1900

%	(α)magn.	%	(α)magn.
	20°		
35.97	1.908	11.83	2.210
26.40	2.004	4.37	2.071
21.10	2.019		

Kohlrausch, 1879

%	κ	$\tau.10^4$
	18°	
2.5	406	228
5	728	229
10	1211	219
20	1626	221
30	1389	229
40	839	285

Hosking, 1904

t	λ 0.00645 N	0.1030 N	0.5203 N	1.0125 N
0	57.0	51.13	43.82	40.03
10	76.8	67.90	58.05	52.42
15	87.1	76.9	65.6	59.1
18	93.1	82.52	70.14	63.26
20	97.2	86.31	73.15	66.33
30	119.8	106.1	89.75	81.41
40	144.2	125.9	107.2	96.78
50	169.5	149.0	125.5	113.1
60	196.6	172.3	144.7	130.1
70	223.6	195.7	164.4	147.7
80	252.6	219.8	184.5	165.3
90	280.4	245.2	204.8	183.2
100	309.8	271.5	225.3	201.9

t	λ 2.937 N	5.02 N	7.36 N	10.71 N
0	29.85	21.23	13.22	5.66
10	39.12	27.60	17.31	7.70
15	44.0	31.0	19.6	8.8
18	46.95	33.15	20.89	9.54
20	49.00	34.53	21.77	10.01
30	59.70	42.08	26.61	12.72
40	70.98	50.1	31.9	15.67
50	82.64	58.2	37.5	18.87
60	94.9	66.8	43.3	22.33
70	107.3	75.7	49.2	25.9
80	119.8	84.8	55.3	29.8
90	132.7	93.6	61.5	33.6
100	145.6	100.0	67.7	37.6

Johnston, 1906

N	λ	N	λ
	100°		
10	46.3	0.25	21.50
9	50.2	0.12	22.92
8	60.9	0.06	24.74
5	97.8	0.025	25.89
2	14.71	0.010	26.50
1	17.26	0.001	31.01

N	λ	N	λ
	0°		
10	7.38	1	35.06
8	14.65	0.5	35.79
6	16.45	0.10	40.43
3	23.08	0.01	43.80
2	26.00	0.001	73.73

Green, 1908

N	κ	N	κ
	17.82°		
0.0928	76.60	5.605	1641.2
0.466	332.0	6.195	1624.1
0.937	600.1	7.040	1555.0
1.530	876.0	8.035	1436.4
2.045	1077.0	8.875	1285.0
2.730	1287.0	9.975	1122.0
3.440	1452.5	11.05	951.8
4.040	1547.0	12.23	785.0
4.875	1624.3	12.42	769.9
	25°		
0.00625	6.86	4.029	1797.0
0.0125	13.34	4.970	1890.2
0.025	21.12	5.325	1904.3
0.0499	49.87	6.218	1885.3
0.1000	96.37	6.976	1824.7
0.2000	182.22	8.445	1607.0
0.4000	338.9	8.804	1544.1
0.6175	492.2	9.173	1476.7
1.1630	825.6	10.400	1247.6
2.001	1233.1	11.265	1100.0
2.860	1537.3	12.345	936.7

Sakhanov, 1913

%	κ	%	κ
	25°		
0	-	35.74	1239
7.331	1156	45.86	717.3
21.28	1896		

Alfimoff, 1917

t	κ	t	κ
	c=60.40		
19.1	620	53.2	1470
30.4	840	63.5	1810
41.4	1130	78.0	2340
	c=33.96		
17.8	1490	60.8	3090
31.4	1930	77.7	3800
47.7	2510		

t	κ	t	κ
	c=20.62		
19.3	1640	65.1	3470
31.2	2080	76.2	3840
50.1	2810		
	c=6.92		
15.3	880	36.1	1320
20.1	970	50.6	1650
29.1	1010	74.4	2180

Jacopetti, 1942

t	κ 0.5 N	1 N	2 N	3 N	4 N
18	352	627	1064	1366	1551
25	407	728	1236	1584	1794
40	532	961	1632	2086	2354
50	625	1126	1911	2438	2776
60	723	1297	2193	2794	3144
80	939	1649	2755	3497	3955
100	1178	2001	3282	4144	4669

t	κ 5 N	6 N	7 N
18	1637	1641	1583
25	1890	1891	1820
40	2469	2461	2363
50	2875	2863	2751
60	3291	3281	3160
80	4131	4147	4045
100	4936	5026	5020

Rodnyanskii and Galinker, 1955

t	λ 1 N	2 N	3 N
25	74.0	62.7	54.0
50	119	96.5	78.5
100	213	162	128
150	295	223	173
200	365	278	214
250	413	316	244
300	430	331	254
340	413	313	238

Kuschel, 1881

%	t	transport number (LiCl)	%	t	transport number (LiCl)
25.46	17.7	0.773	1.00	15.2	0.718
12.74	20.7	0.753	0.50	14.4	0.699
7.30	18.8	0.738	0.20	16.2	0.674
3.38	17.0	0.739			

Scott and Blair jr, 1933

%	χ	%	χ
	29°		
4.12	- 0.716	19.75	-0.703
9.66	0.712	20.85	0.699
11.02	0.710	22.59	0.696
13.88	0.708	26.02	0.689
16.28	0.707	32.80	0.679
17.08	0.706	34.90	0.678
17.97	0.705	42.08	0.669
18.74	0.704	44.49	0.665
18.97	0.705		

Heat constants.

Tucker, 1915

mol% initial	final	Q dil (by mole aq.added)
	15°	
21.45	20.57	960.7
20.57	19.48	858.6
19.48	18.45	767.6
18.45	17.42	644.8
17.42	16.46	554.9
16.46	15.61	460.5
15.61	15.02	391.5
15.02	14.15	330.9
14.15	13.35	282.4
12.53	10.79	172.8
10.79	9.61	136.8
9.61	8.23	82.3
8.23	7.24	61.7
7.24	5.80	40.2
5.80	4.36	31.1
4.36	3.19	21.6

t	Q dil
12.44%	
2.13	221.6
5.98	218.5
22.04	221.5

mol%	U	mol%	U
	20°		
22.74	0.637	8.26	0.815
18.53	0.677	6.00	0.857
13.42	0.736	4.35	0.883
10.81	0.763	0.00	1.000

t	U	t	U
	15.83%		
- 0.55	0.688	35.79	780
+12.65	718	41.19	792
21.67	739		

Jauch, 1921

N	U	N	U
	18°		
0.5	0.9722	3	8584
1	9461	4	8193

Johnson and Molstad, 1951

%	Q vap	%	Q vap
	30°		
45.99	11900	25.11	10639
44.41	11760	24.99	10614
44.40	11760	24.93	10610
44.39	11670	12.41	10520
44.38	11760	12.33	10504
39.89	11510	4.12	10475
39.82	11500	4.00	10482
36.63	11380	3.98	10439
36.55	11360	3.97	10497
31.50	11050	3.95	10421
31.34	11080	3.94	10458
25.37	10659		

%	Q vap	%	Q vap
	50°		
47.47	11820	13.35	10308
47.41	11760	13.23	10291
43.66	11700	13.13	10328
43.58	11660	8.35	10248
39.81	11518	8.29	10240
39.75	11493	4.99	10239
36.95	11237	4.95	10225
36.84	11202	3.26	10234
30.43	10733	3.23	10220
30.29	10702	3.21	10223
22.30	10395	3.18	10244
22.14	10373		

%	Q vap	%	Q vap
	70°		
45.61	11505	28.48	10583
45.71	11540	18.00	10045
36.96	10966	18.13	10069
37.16	10983	18.30	10091
28.35	10515	18.43	10123

Kapustinskii and Ruzavin, 1955

Heat conductivity coefficient $.10^6$ (K)

%	K	%	K
	25°		
3.8	1441	13.2	1400
7.8	1422	24.3	1350

H_2O + Lithium bromide (LiBr)

Heterogeneous equilibria

Tammann, 1885

%	p	%	p
	100°		
0.00	760.0	34.59	511.9
3.58	749.7	40.43	426.3
10.27	724.0	45.61	336.0
15.70	695.7	47.91	295.1
17.45	685.2	50.66	244.7
24.01	635.9	57.59	136.2
31.25	561.6		

Tammann, 1885

t	p				
	0%	15.21%	23.35%	29.89%	35.80%
24.74	23.4	-	18.8	16.0	12.9
33.77	39.4	34.5	30.7	27.7	21.9
39.72	54.5	48.6	43.5	39.3	31.2
46.87	79.1	71.6	64.7	57.9	76.3
50.76	96.1	87.0	78.5	70.3	56.2
53.22	108.4	98.2	88.9	79.5	64.2
55.17	119.1	108.3	98.3	88.4	71.1
59.86	148.5	135.1	123.4	110.3	89.6
63.16	172.7	158.0	144.8	129.6	104.9
65.35	190.6	174.7	159.1	142.3	114.8
66.98	204.9	187.6	170.8	152.7	123.8
68.65	220.5	202.4	183.5	164.3	133.5
71.45	248.9	228.8	208.7	186.4	150.8
73.40	270.5	249.2	226.9	202.9	165.0
74.67	285.3	262.2	238.9	213.9	172.7
77.21	317.1	291.8	266.1	238.0	192.9
78.83	338.9	311.4	283.7	254.6	206.4
80.95	369.4	340.1	309.9	277.3	225.2
82.03	385.7	354.8	323.7	289.4	235.4
83.47	408.5	375.3	342.7	307.0	250.3
86.61	461.9	424.3	387.6	347.2	282.5
89.00	506.3	464.1	424.3	380.3	310.1
90.64	538.8	494.2	451.7	405.0	330.5
92.53	568.5	530.9	485.7	434.6	-
94.75	628.2	576.5	527.3	516.5	-
96.59	672.1	616.5	564.5	506.0	-
100.21	765.7	703.1	642.8	577.5	-

Hüttig and Reuscher, 1924

t	p	t	p
	(2+1) - (1+1)		
4	0.14	26	0.8
17.5	0.6	28	1.0
18	0.6	30	1.2
21.5	0.6	32	2.8
	(1+1) - LiBr		
80	1.8	120	26.1
103.7	9.3	1599	22.8

%	t	p	%	t	p
		(2+1) - sat.sol. - V			
59.2	4	0.14-0.8	62.4	23	1.5
60.3	14	0.8	64.9	30	2.5
60.8	16	0.8	65.8	32	2.8
		(1+1) - sat.sol. - V			
65.8	32	2.8	69.9	70	10.0
67.7	35	2.8	69.9	70	10.5
67.8	40	3.0	70.3	75	12
68.1	45	3.8	70.8	80	14.7
68.4	50	4.8	71.3	85	20.4
68.7	55	5.8	71.8	90	25.8
69.1	60	7.0	72.3	95	30.7
69.4	65.2	9.7	-	159	228

Lannung, 1934

m	p	m	p
	18°		
0.0772	15.03	1.101	4.16
0.0824	15.00	1.250	3.18
0.0904	14.93	1.297	2.65
0.1029	14.90	1.402	2.25
0.1153	14.76	1.446	1.86
0.1393	14.69	1.508	1.75
0.1762	14.33	1.577	1.42
0.2439	13.83	1.662	1.29
0.3368	13.03	sat.d	1.10
0.3959	12.29	1.755	1.03
0.5327	10.92	1.814	0.97
0.674	9.21	1.888	0.79
0.854	6.99	1.901	0.78
0.892	6.64	2.056	0.67
1.006	5.36	2.071	0.58
1.072	4.72		

p dissoc

	18°
(2+1)	0.41
(1+1)	0.006

Johnston 1906

%	b.t.	%	b.t.
0.532	104.14	12.867	137.60
0.791	106.21	14.173	139.02
1.085	108.26	15.597	140.35
1.428	109.87	16.754	141.38
1.798	112.42	17.897	142.64
2.199	114.49	19.092	144.90
2.695	116.57	23.992	145.87
3.171	118.38	27.392	148.25
3.741	120.40	33.392	152.80
4.359	122.29	36.892	154.63
5.765	126.04	41.792	158.25
7.421	129.52	45.392	159.91
9.393	132.82		

Johnston, 1907

%	b.t.	%	b.t.
2.079	100.274	16.57	101.695
4.142	100.532	18.39	103.171
6.228	100.791	20.42	103.741
8.258	101.085	22.31	104.359
9.89	101.428	26.05	105.765
12.43	101.798	29.54	107.921
13.33	102.199		

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1704	0.1657	1.585	1.316
.2197	.2128	1.725	1.419
.2653	.2564	1.917	1.555
.3015	.2863	1.952	1.578
.3794	.3581	1.997	1.610
.4260	.4024	2.395	1.879
.4803	.4489	2.458	1.922
.5212	.4835	2.621	2.023
.5404	.5000	2.644	2.052
.6156	.5653	3.117	2.339
.6201	.5686	3.577	2.624
.6410	.5871	3.917	2.829
.7169	.6520	4.076	2.921
.8886	.7909	4.173	2.974
.9851	.8667	4.216	2.998
1.126	.9799	4.350	3.077
1.150	.9969	4.808	3.325
1.487	1.248		
1.502	1.257		

1 - potassium chloride
2 - lithium bromide

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.943	1.6	1.130
0.2	0.944	1.8	1,163
0.3	0.952	2.0	1.196
0.4	0.960	2.5	1.278
0.5	0.970	3.0	1.364
0.6	0.981	3.5	1.467
0.7	0.993	4.0	1.578
0.8	1.007	4.5	1.687
0.9	1.021	5.0	1.793
1.0	1.035	5.5	1.891
1.2	1.067	6.0	1.989
1.4	1.098		

m	activity coefficient	m	activity coefficient
	25°		
0.1	0.796	1.6	0.917
0.2	0.766	1.8	0.964
0.3	0.756	2.0	1.015
0.4	0.752	2.5	1.161
0.5	0.753	3.0	1.341
0.6	0.758	3.5	1.584
0.7	0.767	4.0	1.897
0.8	0.777	4.5	2.28
0.9	0.789	5.0	2.74
1.0	0.803	5.5	3.27
1.2	0.837	6.0	3.92
1.4	0.874		

Getman and Jones, 1904

N	f.t.	N	f.t.
0	- 0	0.969	- 4.275
0.121	- 0.960	1.940	-10.300
0.242	- 0.905	3.880	-30.500
0.484	- 1.940	4.850	-44.000

Hüttig and Reuscher, 1924

%	f.t.	%	f.t.
59.2 (2+1)	4	68.7	55
60.3	14	69.1	60
60.8	16	69.4	65
62.4	23	69.9	70
64.9	30	70.3	75
65.8 (1+1)	32	70.8	80
67.7	35	71.3	85
67.8	40	71.8	90
68.1	45	72.3	95
68.4	50		

Properties of phases

Kremers, 1858

%	d	%	d
	19.5 %		
0	0.998	30	1.252
5	1.033	35	1.307
10	1.070	40	1.366
15	1.111	45	1.430
20	1.154	50	1.497
25	1.204	55	1.577

t	d			
19.5	1.1233	1.2698	1.3849	1.5452
0	1.1278	1.2768	1.3935	1.5560
40	1.1159	1.2611	1.3751	1.5336
60	1.1066	1.2514	1.3648	1.5222
89	1.0955	1.2405	1.3579	1.5109
100	1.0831	1.2287	1.3423	1.4985

Jones and Getman, 1904

N	d	N	d
	0°		
0	1.005348	0.969	1.117004
0.121	1.014824	1.940	1.235368
0.242	1.027576	3.880	1.286084
0.484	1.057448	4.850	0.999868

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.00174	1.0	1.06051
0.1	1.00490	2.0	1.1215
0.2	1.01109	4.0	1.02976
0.5	1.02976		

Baxter, Boylston and al. 1911

%	d	%	d
	25°		
0	0.9970	3.5637	1.0233
0.1980	0.9984	3.5919	1.0234
0.3328	0.9995	3.7527	1.0245
0.3889	0.9999	4.2994	1.0287
0.4313	1.0001	4.6222	1.0309
0.8136	1.0028	6.1940	1.04335
1.0244	1.0044	14.966	1.1153
1.4355	1.00745	18.190	1.14337
1.6802	1.0092	32.55	1.2889
1.8718	1.0108		

Grufki, 1913

N	d	N	d
	18°		
0	0.99862	1.998	1.12133
0.4992	1.02970	4.000	1.24108
1.000	1.06053		

Lübben, 1913

N	d	N	d
	18°		
0	0.99862	2.146	1.13020
0.500	1.02985	3.999	1.24099
1.025	1.06196		

Gropp, 1915

t	d	t	d
	7.980 N		
0	1.4799	78	1.4415
18	1.4723	100	1.4281
48	1.4580	108	1.4228

Baxter and Wallace, 1916

%	d	%	d
70.19°		50.04°	
64.27	1.77900	64.01	1.78716
41.46	1.37658	41.20	1.38463
20.05	1.13935	19.91	1.14847
10.73	1.08882	10.64	1.06851
5.55	1.01839	5.50	1.02836
2.92	0.99874	2.89	1.00888
25.0°		0.0°	
63.66	1.79812	40.74	1.40237
40.96	1.39392	19.76	1.16381
19.76	1.15763	10.52	1.08245
10.55	1.07766	5.45	1.04134
5.46	1.03742	2.85	1.02135
2.85	1.01795		

de Block, 1925

%	d
16°	
25.5	1.2076
49.3	1.4920
62.2	1.7005

Hüttig and Keller, 1925

N	d	N	d
	20°		
11.235	1.6770	1.15995	1.06349
10.1915	1.6132	1.0068	1.06017
6.27047	1.37558	0.036895	1.03705
4.06685	1.24407	0.25201	1.01392
2.54825	1.15348	0.100162	1.00451
1.54033	1.09267		

Scott, Obenhaus and Wilson, 1934

%	d	%	d
	35°		
60.413	1.72090	31.897	1.17784
54.761	.60856	25.388	.20827
49.949	.52364	17.229	.13097
45.438	.45294	12.959	.09397
43.493	.42456	8.6052	.05845
36.696	.33423		

Scott and Bridger, 1936

%	d	%	d
	35°		
45.183	1.44971	34.706	1.31036
39.999	1.37671	18.076	1.13859

Pohle, 1927

%	π (relative)	%	π (relative)
0	1.000	20.71	810
7.30	0.920	26.80	748
13.51	867	33.11	703

Scott, Obenhaus and Wilson, 1934

%	π	%	π
	35°		
60.413	12.37	31.897	30.94
54.761	24.13	25.388	32.92
49.949	25.69	17.229	35.50
45.438	27.09	12.959	37.17
43.493	27.70	8.6052	38.57
36.696	29.46		

Scott and Bridger, 1936

%	π	%	π
	35°		
45.183	27.12	34.706	30.11
39.999	28.63	18.076	35.28

de Block, 1925

%	σ	%	σ
	16°		
0	73.11	49.3	84.98
29.5	47.51	62.2	89.90

Kremers, 1859

%	n_D	%	n_D
	17°		
0	1.3332	37.58	1.4022
19.61	1.3645		

Borgesius, 1895

%	t	D n_D (sol.-aq.)
17.78	17.9	0.002692
5.17	16.6	0.000678

Jones and Getman, 1904

N	n_D	N	n_D
	0°		
0	1.33395	0.969	1.33810
0.121	1.32711	1.940	1.35062
0.242	1.32869	3.880	1.37455
0.484	1.33191	4.850	-

Baxter, Boylston and al. 1911

%	n_D	%	n_D
	25°		
0	1.33246	3.5637	1.33788
0.1980	1.33273	3.5919	1.33793
0.3328	1.33300	3.7527	1.33825
0.3889	1.33307	4.2994	1.33905
0.4313	1.33309	4.6222	1.33954
0.8136	1.33369	6.1940	1.34200
1.0244	1.33401	14.966	1.35662
1.4355	1.33466	18.190	1.36254
1.6802	1.33499	32.55	1.39194
1.8718	1.33530		

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	2.0	1.35833
0.5	1.33969	4.0	1.38223
1.0	1.34600		

Dn $(H\gamma - H\alpha) \cdot 10^3 = 75$.

Lübben, 1913

N	n 18°		
	2516.0 Å	2558.0 Å	2573.2 Å
0	1.37660	1.37460	1.37361
0.5000	.38676	.38437	.38329
1.025	.39726	.39450	.39345
2.146	.41922	.41564	.41424
3.999	.45476	.44978	.44807
	2748.7 Å	2777.1 Å	2801.0 Å
0	1.36654	1.36570	1.36460
0.5000	.37542	.37448	.37326
1.025	.38457	.38350	.38223
2.146	.40372	.40249	.40095
3.999	.43462	.43296	.43108
	2881.0 Å	2981.1 Å	3018.5 Å
0	1.36260	1.35980	1.35890
0.5000	.37103	.36797	.36698
1.025	.37979	.37636	.37534
2.146	.39796	.39403	.39281
3.999	.42731	.42244	.42089
	3076.0 Å	3081.0 Å	3133.0 Å
0	1.35740	1.35740	1.35622
0.5000	.36536	.36536	.36407
1.025	.37354	.37351	.37219
2.146	.39075	.39066	.38915
3.999	.41837	.41824	.41642
	3255.0 Å	3303.0 Å	3345.6 Å
0	1.35352	1.35200	1.35170
0.5000	.36117	.35957	.35921
1.025	.36901	.36736	.36688
2.146	.38555	.38371	.38314
3.999	.41208	.40996	.40921
	3403.6 Å	3467.0 Å	3611.9 Å
0	1.35062	1.34970	1.34754
0.5000	.35804	.35603	.35473
1.025	.36572	.36459	.36214
2.146	.38176	.38050	.37769
3.999	.40755	.40600	.40265
	4415.1 Å	4678.3 Å	5086.0 Å
0	1.33997	1.33831	1.33633
0.5000	.34672	.34494	.34194
1.025	.35367	.35181	.34966
2.146	.36819	.36612	.36377
3.999	.39147	.38915	.38651

Grufki, 1913

N	n Hα	n Hβ	n Hγ
		18°	
0	1.33141	1.33735	1.34052
0.4992	.33773	.34400	.34740
1.000	.34397	.35056	.35421
1.998	.35606	.36332	.36734
4.000	.37970	.38821	.39360

Hüttig and Keller, 1925

N	n 20°		
	4340 Å	4360 Å	4860 Å
11.235	1.48380	1.48346	1.47638
10.1915	.47169	.47144	.46459
6.27047	.42225	.42199	.41656
4.06685	.39382	.39357	.38894
2.54835	.37427	.37407	.36998
1.54033	.36123	.36103	.35731
1.05995	.35485	.35462	.35112
0.636895	.34928	.34904	.34572
	5460 Å	D	6560 Å
11.235	1.47050	1.46740	1.46368
10.1915	.45893	.45595	.45236
6.27047	.41196	.40937	.40657
4.06685	.38503	.38298	.38035
2.54835	.36652	.36464	.36234
1.54033	.35415	.35244	.35030
1.05995	.34807	.34644	.34450
0.636895	.34282	.34124	.33928

Jones, 1904

and Jones and Getman, 1904

N	λ
0°	
0.121	52.84
0.242	52.07
0.484	46.40
0.969	44.15
1.940	37.30
3.880	29.35
4.850	26.00

Johnston, 1906

N	λ
0°	
6	19.42
3	26.80
2	28.24
1	37.92
0.01	44.04
0.001	73.00

Johnston, 1907

N	λ	N	λ
	100°		
10	4.78	0.5	21.02
7	7.28	0.2	23.27
5	10.02	0.1	24.39
2	15.59	0.001	31.53
1	18.15		

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	87.9	1.0	67.2
0.1	84.4	2.0	57.7
0.2	80.9	4.0	44.2

Gropp, 1915

t	ϰ	t	ϰ
	7.98 N		
-59.1	101.5	-23.4	577.4
-57.8	109.3	-21.8	612.9
-56.15	120.8	-19.3	656.1
-54.55	135.4	-18.4	684.5
-52.8	146.6	-17.45	704.5
-52.3	154.3	-16.45	729.0
-51.7	161.0	-15.35	754.8
-50.2	174.4	-14.4	779.0
-48.8	188.0	-13.65	800.0
-47.7	202.1	-12.0	835.2
-45.8	222.2	-10.6	869.0
-43.7	245.4	- 9.6	913.5
-42.2	264.3	- 8.5	953.0
-40.25	289.4	- 6.7	998.8
-38.75	312.3	- 4.55	1038.0
-37.35	336.0	- 3.05	1076.0
-35.15	365.3	- 1.55	1120.0
-33.55	391.5	- 0.25	1155.0
-31.25	426.0	0	1161.0
-30.2	446.1	+18	1685.0
-29.1	468.2	48	2776.0
-27.85	490.6	78	4000.0
-26.15	523.0	100	4964.0
-24.7	547.6	108	5294.0

Okazaki, 1936

%	Verdet's constant (D).10^5	%	Verdet's constant (D).10^5
	20°		
2.163	1355	16.52	1685
4.149	1396	17.91	1713
7.30	1467	29.80	2061
9.03	1504	40.06	2404
	25°		
4.98	1413	35.09	2234
12.14	1579	47.25	2679
24.05	1886	57.01	3083

Thermal constants

Jauch, 1921.

N	U	N	U
	18°		
0.5	0.9521	3	7622
1	9089	4	6996
2	8314		

Hüttig and Wehling, 1927

%	U	%	U
	3-40°		
100	0.1433(3.2-81)	44.32	0.579
57.96	0.476	42.45	.589
54.86	.491	40.86	.601
51.74	.515	39.14	.616
49.06	.532	37.65	.631
46.72	.552	34.89	.659

Lange and Schwartz, 1928

mol%	U	Q diss
0	0.998	-
0.06	.9961	11652
0.15	.9931	11613
0.24	.9853	11590
0.4	.9723	11546
0.99	.9471	11484
1.4	.9237	11428
1.96	.9013	11386
2.3	.8803	11346
2.91	.8602	11318
3.42	.8417	11270
3.85	.8229	11236
4.76	.7887	11166
5.66	.7579	11094
6.54	.7300	11022
7.41	.7026	10950
8.26	.6779	10874
9.09	.6586	10798
9.91	.6386	10715
10.71	.6201	10632
11.50	.6036	10548
12.28	.5882	10457
13.04	.5742	10361
13.79	.5612	10260
14.53	.5493	10160
15.25	.5383	10044
15.97	.5279	9930
16.67	.5186	9795
17.35	.5099	9665
18.03	.5018	9530
18.70	.4943	9385
19.35	.4871	9233
20.00	.4780	9074
20.63	.4735	8908
21.26	.4672	8740
21.88	.4612	8578
22.48	.4554	8406
23.08	.4498	8236
23.67	.4442	8070
29.24	.4391	7904

H_2O + Lithium iodide (LiI)

Heterogeneous equilibria

Tammann, 1885

%	p	%	p
	100°		
0.00	760.0	45.98	371.6
9.66	637.6	57.16	269.9
14.43	623.2	59.05	132.1
25.91	571.5	59.36	124.5
34.21	510.6	62.71	60.3
37.89	475.6		

t	p				
	0%	9.65%	22.38%	31.85%	33.82%
34.61	41.3	40.0	37.1	33.5	31.6
42.45	63.0	60.9	56.4	51.1	49.2
48.25	84.8	82.4	76.4	69.1	66.7
55.42	120.7	117.1	115.8	99.3	95.9
60.17	150.6	146.2	135.9	124.8	-
64.44	183.1	177.3	165.6	152.9	-
67.53	209.9	203.9	189.9	173.5	-
71.89	253.6	245.9	229.8	207.2	-
75.10	290.5	282.4	263.17	240.9	-
77.91	326.4	317.0	295.8	270.8	-
80.95	369.3	358.4	334.8	307.5	-
84.37	423.2	411.8	386.1	352.4	-
86.43	458.7	445.7	416.4	381.5	-
89.08	508.0	497.9	459.7	421.7	-
92.64	581.0	564.7	527.4	483.9	-
96.12	660.4	641.1	599.2	550.3	-
100.03	760.9	738.0	690.7	634.3	-

Hüttig and Pohle, 1924

t	p	t	p
(3+1) - sat.sol.		(2+1) - sat.sol.	
17	2.8	70.5	7
35	7.2	71.5	6.4
70.5	7	78.0	6.0
		77	4.5

t	p
(1+1) - sat.sol.	
77	4.5
88	7.1
125	7.2
130	6.5

Lannung, 1934

m	p	m	p
	18°		
0.713	15.11	8.82	6.12
1.111	14.83	8.84	5.98
1.639	14.43	9.84	4.01
4.332	11.80	10.39	4.88
4.351	11.71	11.50	3.19
5.781	9.98	sat.d.	2.90

	p dissoc
(3+1) - (2+1)	0.138
(2+1) - (1+1)	0.0089
(1+1) - LiI	0.000016

Hüttig and Pohle, 1924

%	f.t.	%	f.t.
60.2	0	-	77 E
62.1	19	81.2	80
64.1	40	81.4	88
66.7	59	82.6	99
72.4	75 (3+1)	85.3	120
74.1	71.5	88.1	130 (1+1)
74.8	70.5 E	88.7	130
75.4	71.5	89.2	130
78.8	79 (2+1)	89.3	130

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.952	1.0	1.080
0.2	0.966	1.2	1.111
0.3	0.980	1.4	1.143
0.4	0.995	1.6	1.176
0.5	1.008	1.8	1.212
0.6	1.022	2.0	1.250
0.7	1.034	2.5	1.351
0.8	1.049	3.0	1.467
0.9	1.063		

m	activity coefficient	m	activity coefficient
	25°		
0.1	0.815	1.0	0.910
0.2	0.802	1.2	0.955
0.3	0.804	1.4	1.007
0.4	0.813	1.6	1.063
0.5	0.824	1.8	1.127
0.6	0.838	2.0	1.198
0.7	0.852	2.5	1.418
0.8	0.870	3.0	1.715
0.9	0.888		

Johnston, 1906

%	b.t.	%	b.t.
2.58	100.170	41.74	105.826
11.89	100.872	44.91	106.840
18.29	101.481	47.96	108.042
24.99	102.272	50.68	109.218
30.54	103.181	53.01	110.409
33.94	103.846	54.47	111.250
38.09	104.866	57.07	112.730

Jones and Getman 1904, Jones, 1904;

Jones and Bassett, 1905

N	f.t.	N	f.t.
0.080	- 0.296	1.290	- 6.140
0.161	- 0.588	2.580	-16.200
0.322	- 1.218	3.22	-25.000
0.645	- 2.700	5.16	-59.800

Properties of phases

Kremers, 1858 and 1860

%	d	%	d
	19.5°		
0	0.198	35	1.342
5	1.036	40	1.412
10	1.077	45	1.487
15	1.122	50	1.572
20	1.170	55	1.667
25	1.222	60	1.774
30	1.278		

t	d				
19.5	1.16370	1.32480	1.46220	1.60250	1.79410
0	1.16916	1.33362	1.47350	1.61591	-
40	1.15503	1.31368	1.44930	1.58783	1.77728
60	1.14447	1.30153	1.43575	1.57308	1.76091
80	1.13207	1.28811	1.42139	1.55788	1.74457
100	1.11823	1.27373	1.40659	1.54263	1.72843

Kohlrausch, 1879

%	d	%	d
	18°		
5	1.0361	20	1.1643
10	1.0756	25	1.2138
15	1.1180		

Kuschel, 1881

%	t	d
32.03	16.8	1.3055
16.64	19.7	1.1383
8.43	17.3	1.0658
8.30	19.0	1.0640
4.26	19.3	1.0320
0.99	16.0	1.0073
0.50	14.5	1.0040
0.19	14.2	1.0018

Röntgen and Schneider, 1886

%	d
18.0°	
16.63	1.1380
8.49	1.0656

Jones and Getman, 1904 ; Jones, 1904;

Jones and Bassett, 1905

N	d	N	d
	0°		
0.080	1.005704	1.290	1.126904
0.161	1.013928	2.580	1.251016
0.322	1.031252	3.22	1.317168
0.645	1.064520	5.16	1.502264

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.00354	1.0	1.09708
0.1	1.00849	2.0	1.19406
0.2	1.01843	4.0	1.38780
0.5	1.04802		

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	1.0955	1.10590
0.1088	1.00953	2.1798	1.21114
0.2189	1.02068	4.3592	1.42318
0.5488	1.05275		

Baxter, Roylston and al., 1911

%	d	%	d
	25°		
1.1083	1.00509	2.5287	1.01567
1.4231	1.00742	4.9689	1.03405
1.4512	1.00755	12.3103	1.09490
2.1422	1.01282	15.4751	1.12291

Lübben, 1913 and Grufki, 1913

N	d	N	d
	18°		
0.5076	1.04875	1.969	1.19153
0.9705	1.09418	2.900	1.28112

Baxter and Wallace, 1916

%	d	%	d
	70.19°		
62.49	1.79312	18.16	1.12609
59.43	1.72186	10.52	1.05853
40.47	1.38176	9.67	1.05169
37.89	1.34727	5.54	1.01802
19.64	1.13984	4.99	1.01463
	50.04°		
62.21	1.80277	18.01	1.13622
59.15	1.73155	10.43	1.06878
40.19	1.39161	9.59	1.06190
37.62	1.35717	5.54	1.02829
19.48	1.14999	4.95	1.02489
	25°		
61.86	1.81446	17.86	1.14624
58.79	1.74329	10.34	1.07838
39.89	1.40286	9.51	1.07146
37.56	1.36829	5.53	1.03762
19.33	1.16009	4.89	1.03417
	0°		
58.44	1.75490	17.78	1.15255
39.64	1.41274	9.47	1.07618
37.11	1.37778	5.53	1.04150
19.23	1.16668	4.88	1.03797

Hüttig and Keller, 1925

N	d	N	d
	20°		
6.6931	1.65164	0.92835	1.08929
5.3613	1.52126	0.92673	1.08912
3.73075	1.36232	0.59644	1.05691
2.2573	1.21890	0.23775	1.02165
1.4571	1.14092		

de Block, 1925

%	d	%	d
	15°		
0	0.9991	40.0	1.4109
10.2	1.0799	50.4	1.5793
16.6	1.1366	56.8	1.7024
28.5	1.2622		

Gibson, 1935

%	d	%	d
	25°		
0.00	0.9970	33.87	1.3189
9.62	1.0723	46.37	1.5039
25.58	1.2246	53.33	1.6291

Röntgen and Schneider, 1886

%	π relative
17.88°	
16.63	0.888
8.49	0.940

Gibson, 1935

%	π (1-1000 bars)	%	π (1-1000 bars)
	25°		
0.00	39.35	33.87	32.29
9.62	37.50	46.37	29.34
25.58	34.24	53.33	27.57

de Block, 1925

%	σ	%	σ
	15°		
0	73.26	40.0	75.62
10.2	73.49	50.4	76.93
16.6	73.59	56.8	78.03
28.5	73.83		

Kremers, 1859

%	t	n_D
0	17	1.3332
37.07	17	1.4080
57.99	18	1.4836

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.0955	1.35579
0.1088	1.33541	2.1798	1.37796
0.2189	1.33780	4.3593	1.42235
0.5488	1.34456		

Baxter, Roylston and al. 1911

%	n_D	%	n_D
	25°		
1.1083	1.33424	2.5287	1.33641
1.4231	1.33468	4.9689	1.34033
1.4512	1.33473	12.3103	1.35311
2.1422	1.33580	15.475	1.35908
2.3577	1.33620		

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.35383
0.2	1.33740	2.0	1.37428
0.5	1.34357	4.0	1.41506

Grufki, 1913

N	n Hα	n Hβ	n Hγ
	18°		
0	1.33141	1.33735	1.34054
0.5076	1.34175	1.34843	1.35212
0.970	1.35104	1.35840	1.36256
1.969	1.37102	1.37987	1.38497
2.900	1.38960	1.39984	1.40981

Lübben, 1913

N	n 18°		
	2748.7 Å	2881.1 Å	2981.1 Å
0	1.36654	1.36260	1.35980
0.5076	.38678	.38923	.37614
0.9705	.40496	.39612	.39082
1.969	.44407	.43016	.42233
2.900	.48024	.46178	.45162
	3403.6 Å	3467.0 Å	3611.9 Å
0	1.35062	1.34970	1.34754
0.5076	.36413	.36301	.36047
0.9705	.37631	.37499	.37194
1.969	.40242	.40087	.39689
2.900	.42677	.42481	.41976

N	4415.9 Å	4678.3 Å	4800.1 Å	5086.0 Å
0	1.33997	1.33831	1.33766	1:33633
0.5076	.35147	.34955	.34877	.34724
0.9705	.36182	.35979	.35886	.35716
1.9696	.38402	.38155	.38041	.37837
2.900	.40467	.40173	.40019	.39787

Hüttig and Keller, 1925

N	n 20°		
	4340 Å	4360 Å	4560 Å
6.6931	1.49125	1.49075	1.48154
5.3613	.46106	.46062	.45258
3.73075	.42413	.42390	.41736
2.2573	.39123	.39102	.38580
1.4571	.37318	.37299	.36853
0.92835	.36130	.36105	.35712
0.59644	.35393	.35374	.35010
0.23775	.34578	.35558	.34229
0.09864	.34267	.34248	.33933
	5460 Å	5893 Å	6460 Å
6.6931	1.47405	1.47016	1.46558
5.3613	.44605	.44263	.43865
3.73075	.41196	.40918	.40580
2.2573	.38150	.37919	.37645
1.4571	.36480	.36280	.36038
0.92835	.35377	.35198	.34973
0.59644	.34697	.34528	.34973
0.23775	.33941	.33785	.33593
0.09864	.33654	.33503	.33312

Kohlrausch, 1879

%	$\tau.10^4$	$\varkappa$
	18°	
5	219	294
10	216	570
15	212	832
20	207	1087
25	203	1337

Jones and Getman, 1904; Jones, 1904;
Jones and Bassett, 1905

N	λ	N	λ
		0°	
0.080	51.40	1.290	42.50
0.161	49.65	2.580	42.00
0.322	49.50	3.22	34,27
0.645	46.44	5.16	23.63

Johnston, 1906

N	λ	N	λ
		99.4°	
9	61.9	0.5	164.6
7	79.6	0.25	176.7
5	99.8	0.10	191.8
3	125.3	0.001	345.0
1	159.8		

Heydweiller, 1909

N	λ	N	λ
		18°	
0.05	89.4	1.0	69.2
0.1	85.0	2.0	60.6
0.2	81.5	4.0	46.2
0.5	75.4		

Kuschel, 1881

%	t	transport number
32.03	16.8	0.719
16.64	19.7	0.712
8.43	17.3	0.718
8.30	19.0	0.700
4.26	19.3	0.706
0.99	16.0	0.692
0.50	14.5	0.702
0.19	14.2	0.682

Okazaki, 1936

%	Verdet's constant (D).10^4	%	Verdet's constant (D).10^4
		25°	
5.192	1463	30.72	2459
10.99	1660	35.63	2701
14.43	1776	42.61	3091
18.72	1940	48.74	3487
25.03	2196	49.92	3571

Heat constants .

Jauch, 1921

N	U	N	U
		18°	
0.5	0.9310	2	0.7668
1	0.8710	3	0.6772

H_2O + Lithium azide (LiN_3)

Rollet and Wohlgemuth, 1934

%	f.t.	%	f.t.
0	0	38	0 (1+1)
10	-10	40	+19
20	-28	43	40
24	-40	48	68.2(1+1)-LiN_3
26	-47.5 E	50	100 LiN_3
33.5	-31.0 (4+1)-(1+1)		

H_2O + Lithium thiocyanate (LiCNS)

Heterogeneous equilibria

Nikolajev, 1929

mol%	f.t.	mol%	f.t.
83.34	260.0	25.00	18.5
58.13	201.5	24.80	17.6
53.19	144	24.32	11.5
51.55	102	23.81	10.0
50.00	60.5(1+1)	22.53	5.0
47.40	58	21.15	0
45.25	55.5	11.91	- 33.0 E
37.31	44.5	7.04	- 18.0
35.09	38.5	5.50	- 13.7
33.34	34.0	4.86	- 10.8
29.57	27.5	3.09	- 4.0
25.13	19.5	0	0

Iwate , 1930

%	f.t.	%	f.t.
53.2 (2+1)	20	58.8 (1+1)	35
54.5	25	60.4	40
56.7	30		

tr.t. = 34°

Grufki, 1913

N	d	n Hα	n Hβ	n Hγ
		18°		
0	0.99862	1.33141	1.33735	1.34056
0.5003	1.01110	1.33961	1.34600	1.34948
1.024	1.02400	1.34805	1.35493	1.35871
2.043	1.04840	1.36430	1.37205	1.37636
3.436	1.08117	1.38638	1.39536	1.40054

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	2.0	1.35781
0.5	1.33971	9.0	1.36993
1.0	1.34566		

Dn (Hγ -Hα) . 10^3 = 100

Heydweiller and Grube, 1916

w.l.	Dn (sol. - aq.).10^5	
	18°	
	0.5109 N	1.031 N
2749	1238	2485
2981	1124	2260
3256	1045	2101
3405	1014	2042
3612	982	1973
4679	890	1796
	2.082 N	3.465 N
2749	4953	8124
2981	4521	7450
3256	4204	6959
3405	4083	6757
3612	3948	6544
4679	3590	5936

Water + Lithium hydroxide (LiOH)

Ueda, 1933

t	p (1+1)-sat.sol.	p (1+1)-LiOH
20	14.79	2.60
25	19.97	3.90
30	26.81	5.74
35	35.43	8.56
40	46.54	12.41

Korobkov and Galinker, 1956

mol %	P kg	mol %	P kg	mol %	P kg
300°		350°		374°	
28	66	33	98	35	138
24	68	26	113	26	157
20	70	20	120	20	170
14	76	16	131	15	177
7	83	8	146	8	196
3	85	3	161	4	211
400°		420°			
44	178	46	222		
30	206	34	245		
22	221	26	264		
16	234	20	276		
9	250	14	287		
6	259	11	300		

Pickering, 1893

%	f.t.	%	f.t.
13.31	+ 81.1	8.12	--12.1
13.20	+ 78.5	8.07	- 12.5
13.01	+ 77.3	7.70	- 11.45
12.69	+ 71.85	7.40	- 11.5
12.60	+ 72.0	7.23	- 10.5
12.36	+ 66.5	6.70	- 10.6
12.03	+ 58.5	6.66	- 9.7
11.91	+ 58.0	5.99	- 8.7
11.68	+ 45.5	5.86	- 9.5
11.36	+ 33.5	5.44	- 7.8
11.28	+ 10.0	4.99	- 7.5
10.65	- 17.4	4.89	- 6.75
10.38	- 16.75	4.31	- 5.9
10.12	- 16.45	4.05	- 6.0
9.89	- 15.6	3.59	- 4.9
9.53	- 14.9	3.20	- 4.4
9.45	- 14.8	2.83	- 3.8
9.21	- 14.18	2.00	- 2.7
8.82	- 13.25	1.39	- 1.85
8.77	- 13.4	1.16	- 1.5
8.52	- 12.95	0	0

(1+1)

Robinson and Stokes, 1949

m	osmotic coefficient	activity coefficient
	25°	
0.1	0.920	0.760
0.2	0.902	0.702
0.3	0.890	0.665
0.4	0.881	0.638
0.5	0.875	0.617
0.6	0.869	0.599
0.7	0.866	0.585
0.8	0.863	0.573
0.9	0.861	0.563
1.0	0.860	0.554
1.2	0.863	0.542
1.4	0.866	0.532
1.6	0.870	0.525
1.8	0.872	0.518
2.0	0.875	0.513
2.5	0.882	0.503
3.0	0.886	0.494
3.5	0.889	0.487
4.0	0.892	0.481

Okazaki, 1933

%	Verdet's constant (3441 Å)	%	Verdet's constant (3441 Å)
	28°		
2.82	0.04649	9.45	0.05202
5.89	0.04897	10.33	0.05263

Lithium nitrite ($LiNO_2$) + H_2O

Oswald, 1914

%	f.t.	%	f.t.
11.1	- 7.5	50.9	+ 25
15.0	-11.7	55.5	38.5
21.2	-21	56.9	42
22.5	-22.5	60.6	49
29.0	-28.8	63.8	65
30.11	-28.8	68.7	81.5
-	-30.8 E	72.4	91
33.9	-19.3	91.8	96
41.5	0	94.3	92.5
48.9	19	(1+1)	(1+2)

Bureau, 1935

%	f.t.	%	f.t.
0	0	50	+ 18.5
10	- 7.5	55	33
20	-21.0	60	44.5
25	-33.0	63	50.9
26.58	-38.7	65	60
30	-30	70	74
40	-11.5	75	92
44	- 7.95	76	94
(3+2)	(1+2)		

Bureau, 1937

%	f.t.	E	tr.t.
7.01	- 4.9	-38.7	-
12.31	-10.25	-38.6	-
21.08	-24.15	-38.7	-
26.58	-38.7	"	- (3+2)
32.9	-22.7	"	-
40.5	-10.95	"	-
43.5	- 7.95	"	-
45.9	-	"	-
53.50	-	"	- 7.45 (1+1)
59.0	-	"	- 7.95 (1+1)
60	+50.9	-38.6	-50.9 (1+2)
60.75	56.55	"	- 7.95
63	50	"	- 7.95
64.5	-	"	- 7.95 (1+2)
74	92.0	-	- 7.95 and +50.9
75.9	-	-	+50.9 and +94.0 anh.
81.5	-	-	+50.9 and +94.2
88.9	-	-	+94.1

Water (H_2O) + Lithium chlorate ($LiClO_3$)

Heterogeneous equilibria

Kraus, 1908

m	f.t.	m	f.t.
2	-4.42	4.5	-5.265
2.5	-4.493	5	-5.507
3	-4.612	5.5	-5.724
3.5	-4.80	6.0	-5.94
4	-5.30	6.50	-6.16

Berg, 1929

wt%	mol%	f.t.
100	100	127.5
99.5	97.5	126
99.0	95.3	124
97.6	88.8	118.5
97.0	87.0	113.5
95.9	82.3	106.6
95.5	80.8	105
95.4	80.5	104
95.1	79.4	100
95.0	79.0	100
94.5	77.4	95
93.8	75.0	90
93.0	72.7	90
93.0	72.7	89
92.1	70.8	85
91.9	70.0	84
91.0	66.8	82.0
91.0	66.8	80.9
89.2	63.2	70.0
88.8	62.0	68.0
88.1	59.6	60.0
87.4	57.8	55.0
87.1	57.1	55.0
86.1	55.1	50.0
85.7	54.4	45.0
84.9	52.6	42.0
84.7	52.4	40.0
84.2	51.5	35.0
82.9	49.1	25.0
82.5	48.4	20.0
	(1+3)	
82.3	48.1	30.0
81.1	46.1	25.0
80.4	45.0	20.0
	(1+1)	
79.5	43.6	20.0
77.7	41.0	18.0
76.8	39.7	16.2
75.2	37.6	12.7
74.1	36.3	9.2
73.7	35.8	8.0
71.0	32.8	0.0
	(3+1)	
65.6	27.5	7.4
57.4	21.1	6.0
54.1	19.0	3.0
53.1	18.4	0.0

Cavallaro, 1941

m	f.t.	m	f.t.
0.000516	-0.00190	0.36550	-1.3052
0.001015	-0.00373	0.50314	-1.8204
0.002100	-0.00769	0.63021	-2.3091
0.004988	-0.01814	0.76745	-2.8536
0.010121	-0.07318	0.86140	-3.2358
0.020417	-0.07318	1.0311	-3.9378
0.069044	-0.24456	1.2241	-4.803
0.12276	-0.43316	1.5733	-6.438
0.19920	-0.70413	1.8781	-7.782
0.28004	-0.9940	2.1447	-8.904

Properties of phases

Heydweiller, 1912

N	d	N	d
	18°		
0.110	1.00477	1.096	1.0585
0.219	1.01081	2.19	1.1171
0.548	1.02895		

Sakhanov, 1913

%	d	%	d
	25°		
0	0.9971	38.1	1.2919
9.26	1.0578	60.7	1.5554
20.6	1.1400	75.9	1.8000

Oswald, 1914

d = 1.3184 19° 48.9 %

Bureau, 1937

%	t	d
	sat. sol.	
1	44.0	-
18.4	49.6	1.314
34.7	55.85	1.350
56	65.5	1.411
66	66.4	1.429
79.5	70.2	1.452
99	76.4	1.453

Sakhanov, 1913

%	η	%	η
	25°		
0	895	38.1	2118
9.26	1033	60.7	8263
20.6	1276	75.9	53730

Heydweiller and Grube, 1916

w.l. Å	D n (sol. - aq.).10^5 0.5124 N	1.023 N	2.043 N	4.090 N
		18°		
2574	540	1070	2073	3986
2749	526	1041	2016	3873
2981	510	1009	1960	3751
3256	499	987	1915	3674
3405	497	979	1898	3637
3612	488	968	1873	3594
4679	472	934	1809	3468

Heydweiller, 1912

N	κ	N	κ
	18°		
0.110	80.6	1.096	610.2
0.219	152.3	2.19	1003.0
0.548	343.1		

Sakhanov, 1913

%	κ	%	κ
	25°		
0	715.4	38.1	906.0
9.26	1294.0	60.7	401.4
20.6	1538.0		

Jauch, 1921

N	U	N	U
	18°		
0.5	0.9582	3	0.7906
1	.9204	4	0.7360
2	.8517		

Water + Lithium bromate ($LiBrO_3$)

Simons and Waldeck, 1931

%	f.t.	%	f.t.
61.6	5	72.4	53
63.3	15	72.6	56
65.4	25	74.3	70.5
67.5	35	76.2	85
71.5	50	78.0	100

tr.t. anhydre - (1+1) 51°

Averko - Antonovitch, 1943

%	f.t.	%	f.t.
	I		
84.6	143 (b.t.)	67.78	35.0
81.2	121	65.54	24.9
79.6	111	64.51	20.1
78.6	100.5	61.23	0
75.54	80.0	52.0	- 40.0
73.85	65.0	40.0	- 20.0
72.72	55.0	30.6	- 9.8
71.8	50.0	20.3	- 4.8
70.4	45.0	10.3	- 1.05

E : 54.5% - 47°

II

%	f.t.
68.2	17.5
66.8	4.0

%	d	η (water = 1)
	20°	
52.0	1.5845	3.82
	23°	
40.0	1.4048	2.22
30.6	1.2942	1.63
20.3	1.1656	1.31
10.3	1.0751	1.147

Water + Lithium iodate ($LiIO_3$)

Grüneisen, 1905

N	d	N	d
	18°		
3	1.4516	0.05	1.0064
1	1.1737	0.025	1.00264
0.2	1.0298		

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	1.033	1.15823
0.1039	1.01492	2.061	1.31329
0.2072	1.03100	3.084	1.46437
0.518	1.07916		

Heydweiller, 1912

N	d	N	d
	18°		
0.0994	1.01430	1.003	1.15346
0.1005	1.01424	1.5006	1.07635
0.2001	1.02995	1.6080	1.24274
0.2010	1.02976	2.007	1.30448
0.4020	1.06068	2.862	1.43101
0.8040	1.12223	3.084	1.4644

Lühdemann, 1935

N	m	d
	25°	
0	0	0.99707
0.3964	0.4019	1.05846
0.7723	0.7922	1.05846
1.5143	1.5894	1.11617
2.2084	2.3743	1.22818
1.3840	1.4464	1.20859
2.5222	2.7437	1.37803
2.9291	3.2361	1.43788
3.5426	4.0107	1.52757

Grüneisen, 1905

N	η (water=1)
18°	
3	3.272
1	1.3882
0.2	1.0624
0.05	1.0158
0.025	1.00861

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33324	1.0	1.35767
0.1	1.33575	2.0	1.38105
0.2	1.33823	3.0	1.40367
0.5	1.34559		

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.033	1.35850
0.1039	1.33584	2.061	1.38247
0.2072	1.33841	3.084	1.40561
0.518	1.34602		

Lühdemann, 1935

N	n_D	N	n_D
	45°		
0	1.33254	2.5222	1.39151
0.3964	.34221	2.7165	.39599
0.7728	.35120	2.9291	.40073
1.3840	.36253	3.5426	.41428
2.2084	.38449		

Heydweiller, 1912

N	$\varkappa$	N	$\varkappa$
	18°		
0.0994	51.6	0.8040	273.1
.1005	52.0	1.003	314.3
.2010	94.3	1.6080	396.5
.2001	94.5	2.007	428.6
.4020	165.9	2.962	442.0
.5006	199.5	3.084	436.2

Jauch, 1921

N	U	N	U
	18°		
0.5	0.9187	2	0.7566
1	0.8552	3	0.6820

Water + Lithium nitrate ($LiNO_3$)

Heterogeneous eqilibria

Tammann, 1885

t	p 0%	13.73%	25.94%	37.13%	45.51%
30.33	33.2	29.2	24.6	20.2	15.7
38.04	49.8	44.6	37.4	29.6	22.9
46.62	78.1	77.0	60.9	48.3	38.1
50.34	94.1	85.7	73.5	58.3	45.6
55.37	120.2	109.9	94.5	75.4	59.2
59.49	146.0	132.6	134.3	90.8	71.3
63.49	174.7	159.3	137.0	109.3	86.0
67.07	205.7	188.0	161.9	129.2	102.0
73.48	271.4	248.2	213.8	170.8	135.4
76.71	310.7	284.7	244.9	196.0	155.6
80.44	361.9	331.3	286.0	229.3	182.7
82.38	391.2	359.8	308.9	247.3	196.6
87.56	479.2	438.2	398.6	304.1	242.9
92.22	571.9	523.1	452.1	364.6	-
100.66	777.9	710.3	615.5	-	-

Tammann, 1885

%	p	%	p
	100°		
0	760.0	27.59	587.2
2.22	753.2	35.90	503.7
5.14	740.5	42.74	427.7
8.26	726.4	45.63	395.1
17.41	670.3	56.20	256.2
24.56	614.1		

Lincoln and Klein, 1907

%	p	%	p
	25°		
36.38	14.97	14.40	21.86
32.37	16.74	10.40	22.17
28.68	17.69	7.76	22.69
24.14	19.03	4.89	23.18
18.75	20.47	0	23.76

Pearce and Nelson, 1932

m	p	m	p
	25°		
0.0	23.752	3.3106	20.518
0.10058	23.676	4.5706	19.028
0.20179	23.593	5.9276	17.481
0.40594	23.431	7.3940	15.843
0.61264	23.256	8.9836	14.128
0.82194	23.071	10.7055	12.403
1.0339	22.882	11.88331	11.359
2.1115	21.823	12.8639	10.456

Campbell, Fishman and al., 1956

t	p	t	p
	10.00%		
30	30.48	80	335.90
40	52.57	90	497.20
50	87.83	100	718.35
60	141.76	102	771.08
70	221.40		
	20.00%		
30	27.83	70	203.93
40	48.35	80	309.12
50	80.67	90	456.21
60	130.60	100	658.05
		105	784.04
	30.00%		
30	23.08	70	171.32
40	40.22	80	261.16
50	66.79	90	388.41
60	109.42	100	562.25
	40.00%		
30	19.18	80	215.06
40	33.11	90	320.30
50	55.73	100	463.97
60	89.34	105	556.78
70	140.45		
	50.00%		
30	13.59	70	102.21
40	23.37	80	157.20
50	39.93	90	236.10
60	64.66	104	400.57
	57.28%		
30	9.21	80	118.25
40	16.61	90	179.31
50	28.49	100	264.20
60	46.86	105	319.87
70	75.77		
	64.953%		
30	-	80	79.77
40	11.14	90	122.22
50	19.01	100	182.06
60	31.74	105	221.34
70	50.77		

Johnston , 1906

%	b.t.	%	b.t.
0.366	100.103	16.85	102.241
1.09	100.193	20.40	102.918
1.793	100.278	24.84	103.921
2.45	100.264	27.36	104.428
3.19	100.441	32.41	106.160
6.43	100.830	35.95	107.200
7.50	100.841	39.34	108.496
12.34	101.516		

Kremers, 1856

%	f.t.	%	f.t.
32.57	0	66.22	70
43.10	10	69.44	100
62.89	40	71.94	110
62.50	40.5		

sat.sol. b.t.= 200°

Donnan and Burt, 1903

f.t.	%	f.t.	%
0.10	34.8(3+1)	29.86	56.68
10.50	37.9	29.64	57.48
12.10	38.2	29.55	58.03
13.75	39.3	43.6	60.8 (1+2)
19.05	40.4	50.5	61.3
22.10	42.9	55.0	63.0
27.55	47.30	60.6	63.6
29.47	53.67	64.2	64.9 anh.
29.78	55.09	70.9	66.1
29.87	56.42		

E = 17.8° tr.t. = 29.6°

Robinson, 1935

Isopiestic solutions

m_1	m_2	m_1	m_2
	25°		
0.1064	0.1053	1.977	1.707
.1545	.1515	1.989	1.712
.2697	.2627	2.287	1.933
.3816	.3656	2.378	2.009
.4919	.4655	2.665	2.230
.5120	.4846	2.712	2.264
.6345	.5925	2.731	2.276
.7695	.7093	2.809	2.335
.9991	.9070	2.893	2.397
1.082	.9780	3.087	2.550
1.109	1.003	3.130	2.577
1.111	1.003	3.257	2.673
1.211	1.084	3.415	2.791
1.265	1.131	3.524	2.870
1.430	1.260	4.095	3.279
1.456	1.284	4.228	3.378
1.577	1.385	4.616	3.658
1.591	1.392	4.81	3.805
1.647	1.440		

m_1 = molality of potassium chloride
m_2 = " " Lithium nitrate

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.938	1.6	1.052
0.2	0.935	1.8	1.070
0.3	0.940	2.0	1.088
0.4	0.946	2.5	1.134
0.5	0.954	3.0	1.181
0.6	0.962	3.5	1.227
0.7	0.970	4.0	1.270
0.8	0.978	4.5	1.312
0.9	0.987	5.0	1.352
1.0	0.997	5.5	1.387
1.2	1.015	6.0	1.420
1.4	1.033		

Properties of phases

Kremers, 1861

%	d	%	d
	19.5°		
0	0.9983	29.48	1.1947
11.26	1.0671	35.40	1,2428
20.88	1.1317	44.25	1.3172

t			d		
19.5	1.0751	1.1327	1.1910	1.2529	1.3132
0	1.0814	1.1414	1.1987	1.2647	-
40	1.0662	1.1222	1.1764	1.2400	1.2973
60	1.0558	1.1108	1.1670	1.2270	1.2840
80	1.0439	1.0985	1.1541	1.2136	1.2697
100	1.0308	1.0852	1.1406	1.1999	1.2556

Perkin, 1893

%	d	%	d
	15°		
56.56	1.4416	18.17	1.1164
26.16	1.1733	0	0.9991

Chéneveau, 1907

%	d	%	d
	19°		
19.31	1.1234	8.27	1.0495
15.80	1.0986	3.84	1.0224
12.12	1.0739	0	0.9984
10.21	1.0617		

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.00061	1.0	1.03734
0.1	1.00272	2.0	1.07729
0.2	1.00662	4.0	1.15369
0.5	1.01869	5.4	1.2057

Applebey, 1910

m	d	m	d
	25°		
0	0.99707	4.852	1.01666
1.296	1.00225	6.066	1.02161
1.378	1.00262	8.283	1.03080
1.458	1.00292	9.059	1.03401
1.571	1.00338	9.960	1.03741
1.724	1.00396	10.968	1.04205
1.971	1.00500	12.763	1.04967
2.137	1.00566	18.639	1.07528
3.145	1.00969	22.027	1.09047
3.255	1.01008	24.602	1.10223
3.605	1.01157	34.086	1.14759
3.905	1.01276	58.760	1.28342
4.248	1.01325		

Buchanan, 1912

m	d	m	d
	19.5°		
0.008	0.998707	3	1.1063
0.016	0.999025	4	1.1373
0.031	0.999660	5	1.1665
0.062	1.000916	6	1.1940
0.125	1.003395	7	1.2194
0.25	1.008389	8	1.2437
0.5	1.018047	9	1.2663
1	1.037200	10	1.2885
2	1.073000		

Grufki, 1913

N	d	N	d
	18°		
0	0.99862	2.013	1.07804
0.5013	1.01875	4.002	1.15331
1.059	1.04073		

Gropp, 1915

t	d	t	d
	5.939 N		
0	1.2322	78	1.1884
18	1.2242	100	1.1730
48	1.2073		

Ahlmann and Enroos, 1922

t	d 4.38%	8.67%	13.22%	17.44%
10.0	1.0349	1.0686	1.1075	1.1447
12.5	.0344	.0680	.1064	.1434
15.0	.0337	.0671	.1054	.1422
17.5	.0331	.0662	.1042	.1409
20.0	.0324	.0653	.1032	.1396
22.5	.0317	.0644	.1021	.1383
25.0	.0309	.0634	.1009	.1371
27.5	.0301	.0624	.1997	.1357
30.0	.0291	.0614	.1985	.1344
32.5	.0283	.0602	.1972	.1330
35.0	.0273	.0591	.1958	.1315

Pearce and Nelson, 1932

m	d	m	d
	25°		
0.0	0.99707	3.3106	1.11303
0.10058	1.00116	4.5706	1.15094
0.20179	1.00494	5.9276	1.18825
0.40594	1.01295	7.3940	1.22515
0.61264	1.02074	8.9836	1.26187
0.82194	1.02847	10.7055	1.29886
1.0339	1.03616	11.8331	1.32024
2.1115	1.07507	12.8639	1.33885

Guillaume, 1946

%	d
20°	
8.73	1.0524
23.05	1.1512

Campbell, Debus and Kartzmark, 1955

%	d	%	d
	25°		
0.0635	0.997	30.65	1.201
0.6878	1.001	32.37	1.215
3.739	1.018	37.88	1.259
6.384	1.035	40.09	1.277
6.624	1.036	45.46	1.323
12.13	1.070	46.97	1.339
17.13	1.103	47.02	1.340
18.35	1.111	50.33	1.368
21.09	1.130	55.56	1.427
22.39	1.139	57.66	1.456
28.32	1.182	62.36	1.498
	110°		
0.701	0.956	32.45	1.160
3.499	0.973	40.13	1.222
6.41	0.988	45.67	1.270
12.14	1.024	49.00	1.300
17.25	1.053	55.90	1.364
20.49	1.076	60.40	1.405
26.25	1.117	67.10	1.476

Pohle, 1906

%	π relative	%	π relative
	13.9°		
0	1.000	17.01	855
5.90	0.935	22.60	795
10.91	902	27.86	770

Applebey, 1910

m	η (water=1)	m	η (water = 1)
	25°		
		0.5385	1.05975
0.0174	1.0026	0.8666	1.0980
0.0299	1.0040	0.9963	1.1112
0.0567	1.0067	1.316	1.1567
0.0825	1.0099	2.2819	1.3151
0.1071	1.0125	3.8541	1.74075
0.2333	1.0267	4.578	2.0577
0.3238	1.0354	5.849	3.0255
0.3643	1.0405		
	18°		
		1.283	1.14395
0.00724	1.00124	1.475	1.1699
0.0131	1.00200	2.528	1.34985
0.0379	1.0047	2.550	1.3579
0.0784	1.00905	3.120	1.4906
0.1446	1.01545	3.270	1.5367
0.2653	1.0278	4.363	1.9346
0.7034	1.0737		
	0°		
0.0401	1.0032	0.8577	1.0616
0.0833	1.0058	1.134	1.0875
0.1026	1.0076	1.572	1.1345
0.2294	1.01545	2.099	1.2067
0.4179	1.0278	2.508	1.2770
0.4818	1.0325		

Campbell, Debus and Kartzmark, 1955

%	η(water=1)	%	η(water=1)
	25°		
0.0635	1.005	30.65	1.984
0.6878	1.013	32.37	2.122
3.739	1.063	37.88	2.637
6.384	1.109	40.09	2.914
6.624	1.111	45.46	3.849
12.13	1.221	46.97	4.218
17.13	1.351	47.02	4.236
18.35	1.390	50.33	5.142
21.09	1.490	55.56	7.987
22.39	1.534	57.66	-
28.32	1.826	62.36	14.88

%	η(water=1)	τ
	110°	
0.701	1.010	2.950
3.499	1.053	2.953
6.41	1.094	2.987
12.14	1.185	3.020
17.25	1.276	3.127
20.49	1.346	3.218
26.25	1.516	3.364
32.45	1.739	3.710
40.13	2.197	3.956
45.67	2.658	4.427
49.00	3.097	4.585
55.90	4.481	-
60.40	5.48	-
67.10	7.92	-

Walter, 1889

%	n_D	%	n_D
	15°		
0	1.3334	21.5	1.3613
4.01	1.3384	35.5	1.3822
6.42	1.3414	52.3	1.4105
12.2	1.3489		

Chénevau, 1907

%	n_D	%	n_D
	19°		
19.31	1.3582	8.27	1.3433
15.80	1.3533	3.84	1.3378
12.12	1.3484	0	1.3331
10.21	1.3458		

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	2.0	1.34555
0.5	1.33644	4.0	1.35703
1.0	1.33951		

Grufki, 1913

N	n H α	n Hβ	n H γ
		18°	
0	1.33139	1.33733	1.34052
0.5013	1.33558	1.34170	1.34502
1.059	1.34008	1.34631	1.34986
2.013	1.34765	1.35436	1.35806
4.002	1.36267	1.37011	1.37429

Ahlmann and Enroos, 1922

t	n_D 0%	n_D 4.38%	n_D 8.67%
15.0	1.33351	1.34153	1.34915
17.5	1.33323	1.34126	1.34882
20.0	1.33304	1.34104	1.34848
22.5	1.33278	1.34076	1.34822
25.0	1.33257	1.34041	1.34779
27.5	1.33233	1.34010	1.34740
30.0	1.33203	1.33981	1.34711
32.5	1.33181	1.33949	1.34668
35.0	1.33149	1.33914	1.34629

t	n_D 13.22%	n_D 17.44%
15.0	1.35768	1.36570
17.5	1.35729	1.36534
20.0	1.35697	1.36493
22.5	1.35665	1.36449
25.0	1.35629	1.36408
27.5	1.35588	1.36366
30.0	1.35545	1.36323
32.5	1.35512	1.36277
35.0	1.35473	1.36239

Johnston, 1906

N	λ	N	λ
	100°		
6	73.8	0.10	188.6
5	75.8	0.05	196.4
4	85.9	0.025	205.2
2	116.4	0.012	210.5
1	137.4	0.0033	227.2
0.5	154.0	0.0010	258.0
0.25	172.2	0.0005	282.4

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	83.4	1.0	60.5
0.1	79.6	2.0	50.3
0.2	75.2	4.0	34.9
0.5	67.9	5.4	27.25

Gropp, 1915

t	ϰ	t	ϰ
	5.939 N		
- 20.2	523.1	18	1455
- 13.8	655.6	48	2397
- 6.25	829.8	78	3405
0	971.9	100	4141

Campbell, Debus and Kartzmark, 1955

%	ϰ	%	ϰ
	25°		
0.0635	102.5	30.65	1694
0.6878	908.3	32.37	1685
3.739	4244	37.88	1626
6.384	6730	40.09	1578
6.624	6942	45.46	1441
12.13	1112	46.97	1388
17.13	1387	47.02	1387
18.35	1445	50.33	1277
21.09	1545	55.56	1067
22.39	1583	57.66	972
28.32	1685	62.36	817

%	ϰ	τ
	110°	
0.701	2765	2.336
3.499	1114	2.223
6.41	1843	2.166
12.14	2955	2.045
17.25	3649	1.971
20.49	3989	1.960
26.25	4390	1.958
32.45	4476	1.961
40.13	4408	1.974
45.67	4254	2.187
49.00	4127	2.351
55.90	3705	2.687
60.40	3658	3.31
67.10	3317	4.22

Guillaume, 1946

%	(α) magn.10^6	n 5780 Å
	20°	
8.73	3.728	1.3448
23.05	3.297	1.3649

(α) in radians, gauss, centim.

Perkin, 1893

%	(α) magn.	%	(α) magn.
		15°	
56.56	0.8661	18.17	0.9637
26.16	0.9477	0	-

Okazaki, 1933

%	Verdet's constant (3514 A°)	%	Verdet's constant (3514 A°)
		28°	
10.10	0.04099	31.65	0.3939
16.92	.04053	39.37	.3824
23.87	.04004		

Heat constants.

Campbell, Fishman and al. 1956

%	Q vap	%	Q vap
10.00	10211	50.00	10476
20.00	10213	57.28	10844
30.00	10325	64.953	10962
40.00	10310		

H_2O + Lithium perchlorate ($LiClO_4$)

Simmons and Ropp, 1928

%	f.t.	%	f.t.
29.90 (3+1)	0	70.3	92.7
32.88	10	70.33	92.53
35.95	20	70.5	93.2
37.48	25	71.0	97.3
38.87	30	72.8	108.9
41.97	40	75.0	120.7
50.0	64.6	80.0	136.9
55.0	77.9	82.5	144.0
60.0	89.2	85.0	148.5
62.5	92.3	87.5	149.3
65.0	94.3	90.0	144.2
66.32	95.1	91.04	167.5
66.7	94.8	91.11	172
70.0	93.2		

Jones, 1947

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.2267	0.2178	1.9720	1.6448
0.2415	0.2315	2.093	1.734
0.2459	0.2361	2.648	2.144
0.3935	0.3695	3.234	2.561
0.4283	0.4000	3.633	2.841
0.5855	0.5389	4.064	3.149
0.6343	0.5785	4.480	3.439
0.8462	0.7598	4.585	3.513
1.0429	0.9208	5.308	4.038
1.3796	1.1898	5.585	4.230
1.8530	1.5572		

m_1 = m of KCl m_2 = m of LiClO

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.951	1.2	1.104
0.2	0.959	1.4	1.137
0.3	0.971	1.6	1.170
0.4	0.985	1.8	1.204
0.5	0.999	2.0	1.238
0.6	1.013	2.5	1.328
0.7	1.027	3.0	1.419
0.8	1.043	3.5	1.512
0.9	1.058	4.0	1.595
1.0	1.072		

Simmons and Ropp, 1928

t	d	t	d
	sat.sol.		
0	1.215	25	1.269
10	1.236	30	1.277
20	1.258		

Mazzucchelli and Rossi, 1927

%	d 15°	d 25°
4.90	1.02876	1.02610
9.93	1.06056	1.05708
14.91	1.09386	1.08954

Geffcken, 1929

m	d	m	d
	25°		
0	0.99707	2.4877	1.13409
1.9984	1.10968	3.9660	1.20184
2.0625	1.11296	4.8787	1.23970
2.0933	1.11451		

m	n_D	m	n_D
	25°		
0	1.33253	2.4877	1.34716
1.9984	1.344585	3.9660	1.35424
2.0625	1.34486	4.8787	1.35817
2.0933	1.345095		

H_2O + Lithium aminosulfonate (LiH_2NO_3S)

Dawson, Leader and Zimmerman, 1951

m	d 20°	d 10°	d 0°	d -10°
0.822	1.049	1.052	1.053	1.054
1.66	1.089	1.092	1.094	-
2.43	1.123	1.126	1.128	-
3.43	1.168	1.171	1.172	1.173
4.73	1.230	1.233	1.235	1.236

m	η 20°	η 10°	η 0°	η -10°
0.822	1181	1543	2040	-
1.66	1290	1747	2359	-
2.43	1600	2177	2773	-
3.43	1946	2610	3346	3748
4.73	2628	3448	4101	4508

m	κ 20°	κ 10°	κ 0°	κ -10°
0.9	366	315	274	219
1.65	555	417	349	264
2.4	632	492	387	300
3.45	722	585	401	301
4.7	723	586	403	301

H_2O + Lithium sulfite (Li_2SO_3)

Kremers, 1856

b.t. (sat.sol.) = 105°

H_2O + Lithium sulfate (Li_2O_4S)

Heterogeneous equilibria

Tammann, 1885

t	p 0%	p 4.90%	p 7.93%	p 10.08%	p 12.10%
65.00	187.6	183.9	182.1	179.4	177.0
72.35	258.7	253.5	250.9	247.8	245.7
76.28	305.2	300.5	296.9	293.4	290.3
77.81	325.0	319.9	316.2	312.0	309.4
80.38	360.9	354.8	350.8	346.7	343.3
83.94	416.1	409.3	404.8	399.9	396.3
87.39	476.1	468.6	463.1	458.2	453.6
90.25	531.0	522.5	515.9	510.8	505.7
92.71	582.5	573.2	564.7	589.9	555.0

Tammann, 1885

%	p	%	p
	100°		
0	760.0	15.73	722.4
5.65	745.4	18.46	701.5
7.01	744.7	23.20	678.9
9.15	734.3	31.72	65.51

Lescoeur, 1895

t	p dissoc.	p sat.sol.
	(1+1)	
13.4	-	8.6
20	-	12.4
60	22	117
80	61	257
100	184	-
108	264	-

Applebey, Crawford and Gordon, 1934

t	p	t	p
	sat.sol.		
25.30	20.85	89.65	458.0
30.65	28.5	99.50	657.7
45.65	64.5	100.10	676.4
60.15	129.6	105.25	799.0
69.40	197.3	110.35	948.7
84.30	349.9		

Pearce and Eckstrom, 1937

m	p	m	p
	25°		
0.0	23.752	1.0	22.757
0.1	23.651	1.5	22.220
0.2	23.554	2.0	21.646
0.4	23.360	2.5	21.071
0.6	23.164	3.0944 satd	20.363
0.8	22.962		

Campbell, 1944

t	p dissoc.	t	p dissoc.
	(1+1)		
114	339	196	7080
131	741	200	7590
137	891	202	8130
158	1990	207	9330
170	3020	211	11200
177	3890	215	12100
182	4470	221	13800
186	5010	225	16600
188	5370	227	17000
190	5890	230	18600
192	6310	232	19100
194	6760		

t	p	t	p
	sat.sol.		
114	1070	186	7590
131	1860	188	7940
137	2190	190	8520
158	3800	192	8910
170	5370	202	10060
177	6310	224	18100
182	7080	227	18600

Morey and Chen, 1956

t	P kg	t	P kg
	V + L + C		
374	above 1000	500	666
400	430	600	above 1000

Johnston, 1907

%	b.t.	%	b.t.
4.57	100.401	20.52	102.136
7.95	100.700	23.47	102.540
11.46	101.049	25.28	102.777
15.11	101.461	26.56	102.832
17.92	101.487		

Kremers, 1855

%	f.t.	%	f.t.
26.11	0	24.63	45
25.58	20	23.26	65
		22.62	100

Etard, 1894 (fig.)

%	f.t.	%	f.t.
18.4	-20	25.7	- 4
22.5	-16	25.3	+15
22.6	-15	23.9	+90
24.4	-12		

Friend, 1929

%	f.t.	%	f.t.
27.32	-16.0	25.28	38.0
27.21	-13.0	25.12	42.2
27.18	-11.5	25.00	43.7
26.73	- 6.5	24.82	51.6
26.51	+ 0.6	24.71	52.4
26.07	14.0	24.34	65.7
25.96	16.5	24.05	77.0
25.96	16.7	23.76	94.0
25.85	19.6	23.72	103.0
25.47	31.8		

Campbell, 1944

%	f.t.	%	f.t.
4.072	- 1.735	21.95	- 14.65
7.791	- 3.30	24.85	- 18.45
11.30	- 5.11	27.1	- 21.4
14.33	- 7.04	Eut.	- 23.0
17.67	- 9.67		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.818	1.0	0.787
0.2	0.792	1.2	0.800
0.3	0.780	1.4	0.815
0.4	0.775	1.6	0.832
0.5	0.772	1.8	0.847
0.6	0.773	2.0	0.867
0.7	0.775	2.5	0.923
0.8	0.778	3.0	0.984
0.9	0.782		

Properties of phases

Kremers, 1861

%	d	%	d
	19.5°		
0	0.9984	18.43	1.1648
6.89	1.0581	22.72	1.2061
13.27	1.1160		

t	d		
19.5	1.0526	1.0965	1.1760
0	1.0562	1.1012	1.1815
40	1.0457	1.0892	1.1685
60	1.0366	1.0801	1.1597
80	1.0256	1.0694	1.1495
100	1.0130	1.0573	1.1382

Kohlrausch, 1879

%	d
	15°
5	1.0430
10	1.0877

Jahn, 1891

c	%	d
	20°	
23.155	19.75	1.1760
11.946	10.91	1.0941
0	0	0.9982

Forchheimer, 1899

%	d	%	d
	20°		
23.48	1.2330	13.01	1.1133
18.53	1.1650	7.71	1.0658
16.41	1.1449	0	0.9982

Chéneveau, 1907

%	d	%	d
	19°		
18.77	1.1677	8.21	1.0704
15.46	1.1354	4.24	1.0359
11.95	1.1035	0	0.9984

%	t	d
23.32	14.8	1.2112

Tunha, 1914 - 15

N	d 10°	20°	30°	40°	50°
0	0.99973	0.99823	0.99967	0.99114	0.98807
0.1	1.0048	1.0026	1.0002	0.9972	.9920
0.2	1.0091	1.0076	1.0051	1.0013	.9971
0.5	1.0231	1.0213	1.0186	1.0152	1.0107
0.8	1.0363	1.0344	1.0316	1.0281	.0234
1.0	1.0466	1.0437	1.0407	1.0371	.0327
1.5	1.0670	1.0648	1.0617	1.0577	.0533
2.0	1.0880	1.0855	1.0827	1.0784	.0741

Heydweiller, 1921

N	d	N	d
	18°		
0.2	1.00953	3	1.1304
0.5	.02352	4	.1698
1	.04602	5	.2085
2	.0893		

Kohner, 1928

m	d	m	d
	25°		
0	0.99707	1.47010	1.12018
0.379688	1.03167	2.00236	1.15882
1.11587	1.09286	2.68753	1.20480

Applebey, Crawford and Gordon, 1934

%	t	d	
23.40	94.90	1.182	
23.50	100.75	1.179	
23.56	102.5	1.174	(sat.sol.)
23.55	104.0	1.176	

Gibson, 1934

mol %	d	mol %	d
	25°		
0	0.997	15	1.1303
5	1.040	20	1.1774
10	1.0848	25	1.2264

Pearce and Eckstrom, 1937

m	d	m	d
	25°		
0.0	0.997074	1.0	1.083423
1	1.006475	1.5	1.122184
2	1.015627	2.0	1.158514
4	1.033366	2.5	1.192680
6	1.050531	3.0944 satd.	1.231000
8	1.067219		

Tunha, 1914-15

N	η (water t=1) 10°	20°	30°	40°	50°
0.1	1.0348	1.0308	1.0187	1.0248	1.0258
0.2	1.0644	1.0443	1.0405	1.0541	1.0553
0.5	1.1400	1.1211	1.1342	1.1273	1.1230
0.8	1.2275	1.2065	1.1306	1.1103	1.2055
1.0	1.1989	1.2754	1.2982	1.2724	1.2624
1.5	1.4779	1.4533	1.4723	1.4464	1.4259
2.0	1.6917	1.6717	1.6799	1.6276	1.6061

Gibson, 1934

mol%	π (1-1000 bars)	mol%	π (1-1000 bars)
	25°		
0	39.46	15	24.11
5	33.49	20	20.51
10	35.53	25	17.38

Walter, 1889

%	n_D	%	n_D
	15°		
0	1.3334	12.4	1.3542
4.00	1.3401	17.6	1.3630
7.44	1.3458		

Jahn, 1891

c	%	n Hα	n D	n Hβ
		20°		
23.155	19.75	1.3647	1.3667	1.3711
11.946	10.91	1.3499	1.3519	1.3562
0	0	1.3315	1.3322	1.3375

Chéneveau, 1907

%	n C	n D	n Tl	n P	n G'
			14.8°		
23.32	1.37074	1.37276	1.37472	1.37722	1.38069

Heydweiller and Grube, 1916

w.l. A°	D n (.10^5) (sol.-aq.) 0.4307N	0.871 N	1.744N	3.493 N
		18°		
2574	455	880	1669	3092
2749	449	868	1655	3064
2981	445	863	1636	3029
3256	438	849	1617	2998
3405	435	845	1611	2985
3612	430	839	1603	2969
4679	422	826	1576	2921

Kohner, 1928

m	n_D	m	n_D
	25°		
0	1.33253	1.47010	1.35644
0.379688	1.33954	2.00236	1.36337
1.11587	1.35135	2.68753	1.37134

Forcheimer, 1899

%	$(\alpha)_{magn.}$	%	$(\alpha)_{magn.}$
	20°		
23.48	0.392	13.01	0.435
18.53	0.459	7.71	0.509
16.41	0.433		

Jahn, 1891

c	%	(α)magn.
	20°	
23.155	19.75	0.39234
11.946	10.91	0.34874
0	0	1

Kohlrausch, 1879

%	κ	$\tau.10^4$
	18°	
5	399	237
10	608	240

Johnston, 1907

N	λ	N	λ
	99.4°		
5	490	0.5	1283
4	584	0.25	1511
3	698	0.10	1792
2	873	0.001	3265
1	1118		

Next Page

Heydweiller, 1921

N	λ	N	λ
	18°		
0.2	61.1	3	23.26
0.5	50.6	4	18.05
1	41.38	5	13.85
2	30.52		

H_2O + Lithium acid sulfate ($LiHO_4S$)

Tammann, 1885

%	p	%	p
	100°		
0	760.0	17.38	701.7
7.22	740.2	27.26	644.7
13.31	719.6	35.08	584.2

Gillespie and Wasif, 1953

m	κ	m	κ
	25°		
0.1342	176.0	1.0438	533.4
0.3794	327.7	1.2792	568.3
0.6112	424.7	1.5406	595.7
0.8248	487.6	1.7292	608.7

H_2O + Lithium ammonium sulfate (Li_3SbS_4)

Schreinemakers, Cocheret and al., 1907

%	f.t.	%	f.t.
0	0	27.54	-13.11
1.372	- 0.58	29.75	-14.96
2.496	- 0.97	29.84	-15.22
3.560	- 1.35	31.14	-16.23
5.435	- 2.05	32.84	-18.23
7.068	- 2.50	33.98	-19.28
9.101	- 3.20	35.017	-20.15
10.414	- 3.95	35.15	-20.7 E
12.017	- 4.40	35.25	-10
13.524	- 5.12	35.36	0
15.01	- 5.53	35.53	+10
15.83	- 6.00	35.80	+20
17.76	- 6.77	35.87	+30
19.52	- 7.63	35.89	+40
21.205	- 8.56	36.00	+50
22.79	- 9.58	36.05	+60
24.23	-10.52	36.18	+70
26.79	-12.44		

H_2O + Lithium dithionate ($Li_2O_6S_2$)

Tammann, 1885

%	p	%	p
	100°		
0	760.0	27.93	655.3
8.13	734.2	35.29	593.9
15.31	721.2	43.50	505.5
19.83	701.8		

H_2O + Lithium formate ($LiCHO_2$)

Groschuff, 1903 (fig. 1,2,3)

%	f.t.	%	f.t.
21.14.	-20	54.16	91
24.42	0	57.07	98
27.85	+18	57.64	104
35.60	49.5	59.63	120
	(1+1)		

Sidgwick and Gentle, 1922 (fig.)

%	f.t.	%	f.t.
5.80	- 4.27	17.86	- 16.82
10.98	- 9.02	100	+280

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.01285	3	1.0731
1	1.02535	4	1.0963"
2	1.0495		

N	λ	N	λ
	18°		
0.5	48.08	3	19.9
1	38.84	4	14.02
2	27.21		

Previous Page

H_2O + Lithium acetate ($LiC_2H_3O_2$)

Sidgwick and Gentle, 1922

%	f.t.	%	f.t.
2.63	- 1.54	56.60	55.4
4.83	- 2.98	61.92	57.2
9.52	-6.62	-	57.8
18.33	-16.12	64.25	57.8
23.76	0	64.88	57.0
31.28	25.8	66.73	102.8
38.25	36.7	71.33	157.5
49.55	50.5	100	286

(2+1)

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1263	0.1252	2.015	1.875
.1621	.1598	2.365	2.107
.2118	.2082	2.479	2.199
.3397	.3319	2.527	2.239
.5922	.5678	2.565	2.272
.7327	.6951	2.839	2.499
.8090	.7665	2.896	2.545
1.017	.9490	3.113	2.719
1.103	1.028	3.253	2.828
1.232	1.145	3.821	3.294
1.241	1.147	4.067	3.501
1.565	1.431	4.480	3.840
1.741	1.579	4.516	3.869
1.793	1.622	4.810	4.091

m_1 = molality of Potassium chloride
m_2 = " " Lithium acetate

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.935	1.2	0.975
.2	.928	.4	0.988
.3	.929	.6	1.001
.4	.931	.8	.014
.5	.935	2.0	.027
.6	.940	2.5	.061
.7	.945	3.0	.093
.8	.951	3.5	.123
.9	.956	4.0	.153
1.0	.962		

Rubien, 1911 and Heydweiller, 1912

N	d	N	d
	18°		
0	0.99862	1.006	1.02368
0.0998	1.00115	2.078	1.04976
0.1996	1.00369	3.947	1.09381
0.5020	1.01127	4.146	1.09859

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34312
0.1	1.33428	2.0	1.35271
0.2	1.33528	3.0	-
0.5	1.33824	4.0	1.37119

$Dn\ (H\gamma - H\alpha)\ .\ 10^3 = 26$

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.006	1.34318
0.0998	.33427	2.078	.35343
.1996	.33528	4.146	.37247
.5020	.33826		

Heydweiller and Grube, 1916

w.l. A°	Dn (.10^5) (sol.aq.) 0.4885 N	1.047 N	2.022 N	4.033 N
		18°		
2574	621	1314	2471	4842
2749	590	1239	2347	4596
2981	563	1190	2255	4431
3256	545	1153	2180	4294
3405	538	1137	2151	4242
3612	529	1121	2128	4181
3679	507	1073	1930	3994

Heydweiller, 1912

N	κ	N	κ
	18°		
0.0998	51.2	2.078	365.9
0.1996	92.5	3.947	291.4
0.5020	189.2	4.146	277.9
1.006	290.4		

H_2O + Lithium palmitate ($LiC_{16}H_{31}O_2$)

Vold, 1943

mol %	f.t.	mol %	f.t.
100	224	7.7	221
42.5	187 L+s+c	6.5	221 L+s+m
9.90	240 L+s	2.01	138 L+m+so

s = "superneat" c = concentrated form

m = middle soap so = solid soap

H_2O + Lithium lactate ($LiC_3H_5O_3$)

Hoppe, Seyler and Araki, 1895

c	t	d	t	(α)D
		1 (-)		
23.9650	25.0	1.09295	25.8	- 8.998
11.5878	22.0	.04771	22.0	-11.486
9.9946	21.4	.04294	21.4	-10.953
9.5458	15.5	.03872	12.2	-11.209
7.6092	22.6	.03184	22.4	-12.277
5.1268	22.5	.02251	20.4	-12.162
4.9841	14.5	.02085	11.6	-11.912
		d (+)		
9.1203	17.3	1.03677	12.8	+11.905
7.0897	23.0	.02967	21.1	+13.041
5.0997	17.6	.02120	12.8	+12.940
3.7140	22.5	.01609	21.4	+12.655

H_2O + Lithium acid malate ($LiC_4H_5O_5$)

Schneider, 1881

%	d	(α) D
	20°	
49.882	1.2821	-4.64
39.907	1.2181	-6.15
29.956	1.1592	-7.29
19.952	1.1028	-7.94
9.947	1.0493	-8.85

H_2O + Lithium malate ($Li_2C_4O_5H_4$)

Schneider, 1881

%	d	(α) D
	20°	
39.242	1.2493	- 4.10
32.993	1.2074	- 6.12
27.406	1.1720	- 7.63
20.532	1.1268	- 9.02
14.586	1.0886	-10.21
8.380	1.0596	-11.57
6.113	1.0430	-11.95
0	0.9982	-

H_2O + Lithium tartrate ($Li_2C_4H_4O_6$)

Campbell and Slotin, 1933

%	f.t.	%	f.t.
22.74	60.0	21.26	20.0
21.53	45.0	23.94	10.5
21.00	30.0	25.12	8.0
21.07	25.0	29.62	0.0

Peyches, 1936

mol %	(α)Hg g	mol %	(α)Hg g
		20°	
0.004	46.6	0.500	45.6
0.010	46.7	0.750	44.7
0.020	46.8	1.000	43.8
0.050	46.8	1.250	42.8
0.100	46.8	1.500	41.7
0.200	46.6	1.200	40.0
0.300	46.3		

H_2O + Lithium benzoate ($LiC_7H_5O_2$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
7.22	- 1.69	32.20	34.5
13.69	- 3.94	34.61	84.5
19.85	- 6.49	36.51	111.0 0aq.
27.97	- 0	45.14	162.4
29.80	+13.5	49.12	176.0
32.12	28.5		

H_2O + Lithium salicylate ($LiC_7H_5O_3$)

Schlamp, 1894

%	b.t.	%	b.t.
0	100.000	8.66	100.589
1.67	100.129	11.54	100.799
3.35	100.232	13.20	100.923
5.83	100.389	25.51	101.104

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
10.17	- 2.26	56.50	28.5
20.62	- 5.56	57.50	32.0
35.83	-12.82	59.67	38.5
45.25	- 8.5 (6+1)	64.18	52.0
49.04	- 1.0	66.56	60.0
51.96	+ 3.5	66.56	73.0 0 aq.
52.45	+ 9.0 (1+1)	71.46	113.0
52.96	+10.0	75.77	138.0

H_2O + Lithium m-hydroxybenzoate ($LiC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
16.0	- 4.41	55.04	74.5
29.98	-10.78	58.47	104.0
39.97	-17.67	61.86	122.0
52.53	+10.0		

H_2O + Lithium p-hydroxybenzoate ($LiC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
5.04	- 0.98	31.54	-12.62
9.81	- 2.37	30.84	+12.5
17.88	- 5.23	31.96	85.0
25.96	- 9.28	35.00	113.0

H_2O + Lithium methanesulfonate ($LiCH_3O_3S$)

Dawson, Leader and Zimmerman, 1951

m	d				
	20°	10°	0°	-10°	-15°
0.023	0.9952	0.9997	0.9992	-	-
0.062	0.9999	1.002	1.002	-	-
0.124	1.002	1.005	1.006	-	-
0.560	1.027	1.029	1.030	-	-
0.755	1.036	1.038	1.039	-	-
1.211	1.056	1.058	1.059	-	-
1.536	1.069	1.071	1.074	-	-
2.018	1.088	1.090	1.092	-	-
2.714	1.113	1.115	1.118	1.121	-
3.361	1.135	1.137	1.140	1.143	1.145
4.094	1.157	1.161	1.164	1.167	1.169
4.614	1.172	1.175	1.179	1.183	1.187 (-20°)

m	η				
	20°	10°	0°	-10°	-15°
0.023	960	1281	1751	-	-
0.062	1040	1375	1830	-	-
0.124	1169	1464	1902	-	-
0.560	1218	1636	2147	-	-
0.755	1236	1705	2242	-	-
1.211	1284	1763	2253	-	-
1.536	1459	1902	2616	-	-
2.018	1687	2121	2815	-	-
2.714	1967	2572	3552	4760	-
3.361	2333	2952	4285	5620	5850
4.094	2776	3741	5134	6970	7140
4.614	3145	4234	6276	7910	9250 (-20°)

m	κ				
	20°	10°	0°	- 10°	- 20°
0.15	78.5	61.5	45.7	-	-
0.35	253	199	148	-	-
0.75	313	246	184	-	-
1.25	423	327	248	-	-
1.50	465	377	287	-	-
2.00	526	417	314	-	-
2.75	573	451	341	247	-
3.35	585	449	352	245	164
4.10	570	435	326	237	159
4.60	498	387	301	223	-

H_2O + Lithium trichloracetate ($LiC_2O_2Cl_3$)

Schreiner, 1928

N	d	n_C	n_D	n_F
		18°		
0	-	1.33125	1.33311	1.33725
0.494	1.04256	.34207	.34398	.34834
0.994	1.08666	.35294	.35492	.35947
1.470	1.12846	.36330	.36547	.37010

H_2O + Lithium aminosulfonate (LiH_2NO_3S)

Dawson, Leader and Zimmerman, 1951

m	d			
	20°	10°	0°	- 10°
0.822	1.049	1.052	1.053	1.054
1.66	1.089	1.092	1.094	-
2.43	1.123	1.126	1.128	-
3.43	1.168	1.171	1.171	1.173
4.73	1.230	1.233	1.235	1.236

m	η			
	20°	10°	0°	-10°
0.822	1181	1543	2040	-
1.660	1290	1747	2359	-
2.430	1600	2177	2773	-
3.430	1946	2610	3346	3748
4.730	2628	3448	4101	4508

m	κ			
	20°	10°	0°	- 10°
0.9	366	315	274	219
1.65	555	417	349	264
2.4	632	492	387	300
3.45	722	585	401	301
4.7	723	586	403	301

H_2O + Lithium dodecylsulfonate ($LiC_{12}H_{25}O_3S$)

Vold, 1941

%	tr.t. 1	2	3
49	192	33	-
46.6	134	33	117
8.08	-	32	-
4.07	-	30	-
1.93	-	27	-
1.11	-	25	-

1 : liq.cryst. + L

2 : C + L or C + liq.cryst.

3 : f.t.

H_2O + Lithium p - Toluenesulfonate ($LiC_7H_7O_3S$)

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1052	0.1050	1.630	1.639
.2083	.2079	1.687	1.707
.2844	.2823	1.858	1.894
.3340	.3323	1.950	1.988
.4944	.4885	2.107	2.155
.5550	.5550	2.618	2.676
.7120	.7056	2.887	2.972
.7217	.7174	3.355	3.427
1.132	1.123	3.451	3.524
1.240	1.237	3.839	3.899
1.389	1.383	4.483	4.513
1.420	1.419	4.810	4.819
1.463	1.471	4.810	4.822
1.543	1.547		

m_1 = molality of Potassium chloride

m_2 = " " Lithium p-toluenesulfonate

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.928	1.2	0.904
.2	.917	.4	.902
.3	.912	.6	.899
.4	.908	.8	.894
.5	.906	2.0	.893
.6	.906	2.5	.899
.7	.905	3.0	.912
.8	.905	3.5	.930
.9	.905	4.0	.951
1.0	.905	4.5	.972

H_2O + Rubidium fluoride (Rb F)

Lannung, 1934

m	p	m	p
	18°		
0.4176	15.15	12.91	6.09
0.769	15.10	18.37	3.46
1.296	14.80	22.75	2.24
1.672	14.59	26.59	1.53
2.591	14.03	sat.d.	1.22
4.847	12.41	p dissoc.	(1+1) 0.11

Heydweiller, 1921

N	d	N	d
	18°		
0.2	1.01828	2	1.1764
0.5	1.04515	3	1.2609
1	1.0896	4	1.3458

Grufki, 1913

N	d	n Hα	Hβ	Hγ
		18°		
0	1.33139	1.33759	1.34051	0.99862
0.3432	1.33364	1.33958	1.34278	1.02985
1.039	1.33765	1.34365	1.34688	1.09148
2.074	1.34296	1.34896	1.35218	1.18114
4.221	1.35200	1.35801	1.36130	1.36094

Heydweiller, 1921

N	λ	N	λ
	18°		
0.2	92.9	2	71.4
0.5	86.2	3	63.7
1	80	4	57.6

H_2O + Rubidium chloride (Rb Cl)

Heterogeneous equilibria

Tammann, 1885

t	p 0%	13.10%	23.89%	30.18 %	43.44%
34.13	40.2	38.8	36.5	35.7	31.3
36.97	47.0	45.3	42.6	40.7	36.4
40.55	57.0	54.4	51.3	49.7	44.1
43.46	66.4	63.2	59.6	57.6	51.2
47.09	80.0	77.0	72.5	69.5	62.1
49.70	91.2	87.7	82.2	79.39	70.3
52.74	105.9	101.4	95.0	101.3	81.2
54.77	116.8	112.1	104.9	116.0	89.5
57.53	133.2	127.9	120.3	129.6	102.9
59.90	148.8	143.1	134.4	147.4	115.0
62.72	169.3	161.9	152.4	162.8	130.9
64.97	187.4	180.6	169.8	189.0	145.2
68.19	216.1	207.6	195.4	212.1	168.6
70.86	242.6	233.0	219.8	247.9	188.0
74.57	284.2	273.1	257.3	275.4	220.4
77.06	315.2	302.6	285.8	295.1	243.6
78.76	337.4	324.1	305.1	315.1	261.6
80.33	360.2	346.8	326.2	336.9	-
82.00	385.2	370.9	349.0	372.2	299.3
83.81	414.0	398.5	374.8	391.6	321.7
85.85	448.3	431.1	405.6	414.7	337.5
87.83	473.1	455.7	428.7	440.3	367.5
88.80	502.5	484.5	456.1	512.2	389.2
92.82	594.9	563.8	530.6	574.0	455.2
95.85	654.2	630.4	594.1	-	508.5
97.52	695.1	668.3	630.0	609.6	540.9
100.31	768.5	737.8	695.2	673.2	598.7

Tammann, 1885

%	p	%	p
	100°		
12.17	730.9	29.41	671.0
16.52	719.2	33.68	653.3
22.09	700.8	34.23	651.6
28.06	677.5	54.85	511.1

Lannung, 1934

m	p	m	p
	18°		
0.3900	15.29	1.166	14.92
0.4115	15.27	1.288	14.86
0.6192	15.17	1.717	14.67
0.6980	15.14	1.982	14.54
0.795	15.09	3.100	13.99
0.867	15.57	4.310	13.41
0.966	15.03	6.173	12.49
1.098	14.97	sat.d.	11.86

Morey and Chen, 1956

t	P kg
V + L + C	
400	97
500	119
600	115

Buchanan, 1899

%	b.t.	%	b.t.
	747.1 mm		
37.46	104.56	11.69	100.51
32.25	103.55	10.81	100.41
27.69	102.54	9.73	100.31
24.52	102.02	8.59	100.21
20.72	101.52	7.27	100.11
19.42	101.32	5.98	100.01
17.77	101.12	4.65	99.90
15.90	100.92	0.00	99.52
14.05	100.79		

Berkeley and Applebey, 1911

Sat. sol. b.t. = 113.710°

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.923	1.4	0.888
0.2	0.907	1.6	0.890
0.3	0.898	1.8	0.893
0.4	0.893	2.0	0.896
0.5	0.889	2.5	0.905
0.6	0.887	3.0	0.916
0.7	0.886	3.5	0.928
0.8	0.886	4.0	0.941
0.9	0.885	4.5	0.952
1.0	0.885	5.0	0.966
1.2	0.886		

Rimbach, 1902

%	f.t.	%	f.t.
43.61	0.4	46.56	15.5
43.31	1.0	53.71	57.3
45.33	7.0	59.48	114.9

Berkeley, 1904

%	f.t.	%	f.t.
43.61	0.55	53.62	60.25
47.46	18.70	55.46	75.15
49.65	31.50	57.01	89.35
51.52	44.70	59.44	114.0 b.t.

Mariani and Di Giacomo-Dejak, 1950

N	f.t.	N	f.t.
0.017790	- 0.06510	0.603903	1.988
0.042046	0.15115	0.990213	3.244
0.062081	0.21900	1.154962	3.789
0.077613	0.2774	1.248118	4.128
0.078070	0.2738	1.492670	4.947
0.121034	0.4165	1.677191	5.606
0.157323	0.5352	1.895603	6.388
0.186214	0.6296	2.052059	6.977
0.235052	0.7872	2.235987	7.645
0.244901	0.8211	2.534901	- 8.672
0.359231	1.194		

Properties of phases

Berkeley, 1904

%	t	d
43.61	0.55	1.4409
47.46	18.70	1.4865
49.65	31.50	1.5118
51.52	44.70	1.5348
53.62	60.25	1.5558
55.46	75.15	1.5746
57.01	89.35	1.5905
59.44	114.0 b.t.	1.6148

Heydweiller, 1909

N	d	N	d
	18°		
0.1	1.00753	1.0	1.0860
0.2	1.01639	2.0	1.1710
0.5	1.04271		

Schneider, 1910

N	d	N	d
	18°		
2.0	1.1713	0.2	1.01677
1.0	1.0861	0.1	1,00772
0.5	1.04277	0	0.99862

Rubien, 1911 and Heydweiller, 1912

N	d	N	d
	18°		
0	0.99862	0.9920	1.08556
0.0992	1.00749	1.997	1.17131
0.2009	1.01651	4.206	1.3576
0.5016	1.04295		

Buchanan, 1912

m	d	m	d
	19.5°		
0.5	1.0409	5.0	1.3519
1.0	1.0814	6.0	1.4061
2.0	1.1573	7.0	1.4562
3.0	1.2263	7.767	1.4971
			(22.9°)

N	d	N	d
	19.5°		
0.5	1.041446	0.016	0.99977
0.25	1.020204	0.008	0.999078
0.125	1.009377	0.004	0.998721
0.062	1.003903	0.002	0.998535
0.031	1.001139	0.000	0.997931

Grufki, 1913

N	d	N	d
	18°		
0	0.99862	2.074	1.17798
0.5165	1.04422	4.000	1.33815
1.009	1.08699		

Clausen, 1914

t	d			
	0.5123 N	1.001 N	2.073 N	3.984 N
6	1.04607	1.08916	1.18197	1.34331
18	1.04399	1.08671	1.17819	1.33823
30	1.04066	1.08310	1.17386	1.33307

N at 18°

Baxter and Wallace, 1916

t	d	t	d
	50.04°		
51.51	1.53052	7.50	1.4425
31.13	1.26357	3.83	1.01619
14.37	1.10014		
	25°		
48.88	1.50625	9.16	1.06759
42.48	1.41647	6.97	1.05000
36.53	1.33989	5.76	1.04047
29.23	1.25585	4.72	1.03232
24.83	1.20990	3.55	1.02354
20.84	1.17076	2.92	1.01875
13.34	1.10260	2.39	1.01471
	0°		
30.71	1.28165	7.42	1.05840
14.19	1.11611	3.79	1.02923

Geffcken, 1929

m	d	m	d
0	0.99702	3.0162	1.22936
0.8289	1.06723	5.6678	1.39143
1.4150	1.11366		

Pohl, 1852

%	π relative	%	π relative
	13.9°		
0	1.000	26.60	0.735
9.90	0.898	33.77	0.676
17.80	0.845		

Schneider, 1910

N	η (water=1)	N	η (water=1)
	18°		
2.0	0.9405	0.2	0.9915
1.0	0.9645	0.1	0.9969
0.5	0.9790	0	1

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34398
0.1	1.33539	2.0	1.35383
0.2	1.33550	4.0	1.37265

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.9920	1.34390
0.0992	1.33438	1.997	1.35380
0.4009	1.33551	4.206	1.37446
0.5016	1.33875		

Grufki, 1913

N	n		
	H α	H β	H γ
		18°	
0	1.33141	1.33742	1.34051
0.5165	1.33698	1.34311	1.34656
1.009	1.34210	1.34836	1.35172
2.074	1.35283	1.35917	1.36281
4.000	1.37003	1.37722	1.38125

Geffcken, 1929

m	n_{He}	m	n_{He}
	25°		
0	1.33259	3.0162	1.35967
0.8289	1.34109	5.6678	1.37725
1.4150	1.34651		

Heydweiller, 1909

N	λ	N	λ
	18°		
0.1	115.3	1.0	101.7
0.2	111.1	2.0	97.4
0.5	105.3		

Heydweiller, 1912

N	κ	N	κ
	18°		
0.0992	113.1	1.997	1939
0.2009	223.0	2.977	2761
0.5016	529.1	4.206	3642

Clausen, 1914

t	λ			
	0.5123 N	1.001 N	2.073 N	3.984 N
6	81.60	79.88	77.23	74.17
18	105.2	101.0	97.0	87.34
30	130.4	124.1	116.6	108.8

N at 18°

Okazaki, 1933

%	Verdet's constant (3441 Å)	%	Verdet's constant (3441 Å)
	28°		
7.21	0.04615	25.92	0.05378
13.38	0.04882	32.00	0.05687
19.95	0.05134	39.47	0.06031

Okazaki, 1942

%	Verdet's constant (D)	%	Verdet's constant (D)
	25°		
11.90	0.01411	31.27	0.01616
16.18	0.01455	36.12	0.01669
21.26	0.01503	40.71	0.01732
25.91	0.01556		

Kapustinskii and Ruzavin, 1955

Thermic conductivity coefficient $h.10^6$

%	h	%	h
	25°		
3.91	1438	17.6	1369
7.82	1421	23.4	1333
12.7	1396	31.7	1285

H_2O + Rubidium bromide (RbBr)

Heterogeneous equlibria

Lannung, 1934

m	p	m	p
	18°		
0.2832	15.35	0.804	15.11
0.3006	15.33	1.116	14.97
0.3210	15.32	1.308	14.87
0.3289	15.31	1.446	14.82
0.3528	15.30	1.842	14.63
0.3525	15.30	2.366	14.35
0.3843	15.29	3.382	13.88
0.4357	15.28	4.203	13.48
0.5871	15.22	4.778	13.20
0.5959	15.20	sat.d.	12.53

Rimbach, 1905

%	f.t.	%	f.t.
47.26	0.5	56.87	39.7
49.50	5.0	60.39	57.5
51.17	16.0	67.24	113.5

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.922	1.4	0.881
0.2	0.905	1.6	0.882
0.3	0.897	1.8	0.884
0.4	0.892	2.0	0.887
0.5	0.888	2.5	0.893
0.6	0.886	3.0	0.899
0.7	0.884	3.5	0.907
0.8	0.882	4.0	0.916
0.9	0.881	4.5	0.924
1.0	0.881	5.0	0.934
1.2	0.880		

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.2692	0.2712	1.888	1.933
0.5738	0.5813	2.202	2.263
0.8313	0.8441	2.465	2.540
0.8776	0.8936	2.708	2.789
0.8798	0.8930	3.047	3.155
1.030	1.048	3.687	3.854
1.338	1.364	4.213	4.413
1.388	1.419	4.457	4.688
1.766	1.809	4.518	4.758
1.764	1.810	4.612	4.861

m_1=molality of KCl m_2=molality of RbBr

Properties of phases

Rubien, 1911 and Heydweiller, 1912

N	d	N	d
	18°		
0.0998	1.01128	1.010	1.12527
0.2000	1.02402	2.009	1.24866
0.5006	1.06172	4.000	1.49108

Buchanan, 1912

m	d	m	d	
	19.5°			
0.5	1.0596	4	1.4155	
1	1.1175	5	1.4985	
2	1.2261	6=	1.5746	
3	1.3250	6.7229	1.6252	23°

N	d 19.5°	d 23.0°
0.50	1.059519	-
0.25	1.029411	-
0.125	1.013935	1.013316
0.062	1.006227	1.005490
0.031	1.002310	1.001597
0.016	1.000326	0.998598
0.008	0.999354	-
0.004	0.998828	-
0.002	0.998605	-
0.001	0.998450	-
0.000	0.997931	-

Grufki, 1913

N	d	N	d
	18°		
0	0.99862	2.031	1.25133
0.5156	1.06359	4.072	1.50001
1.020	1.12645		

Lübben, 1913

N	d	N	d
	18°		
0	0.99862	2.031	1.25133
0.5108	1.06298	4.072	1.50001
1.020	1.12645		

Clausen, 1914

t	d 0.508 N	1.020 N	2.031 N	4.072 N
		18°		
6	1.06445	1.12936	1.25646	1.50570
18	1.06250	1.12659	1.25224	1.49962
30	1.05888	1.12186	1.24741	1.49414

Baxter and Wallace, 1916

%	d	%	d
50.04°		25°	
52.91	1.61724	52.55	1.62932
32.69	1.30487	32.45	1.35593
15.32	1.11598	15.19	1.12608
8.09	1.05496	8.02	1.06156
4.16	1.01996	4.11	1.02925
0°		0°	
32.25	1.32467	7.99	1.06605
15.12	1.13207	4.10	1.03297

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34784
0.1	1.33476	2.0	1.36177
0.5	1.34064	4.0	1.38840

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.010	1.34799
0.0998	1.33476	2.009	1.36190
0.2000	1.33624	4.000	1.38840
0.5006	1.34056		

Lübben, 1913

N	n 2516.0 Å	2558.0 Å	2573.2 Å	2748.7 Å
		18°		
0	1.37660	1.37460	1.37361	1.36654
0.5108	1.38786	1.38548	1.38439	1.37646
1.020	1.39894	1.39618	1.39500	1.38621
2.031	1.42039	1.41698	1.41557	1.40511
4.072	1.46181	1.45681	1.45513	1.44128

N	n 27771.0	2801.0	2881.1	2981.1
		18°		
0	1.36570	1.36460	1.36260	1.35980
0.5108	1.37552	1.37432	1.37208	1.36901
1.020	1.37520	1.38382	1.38137	1.37800
2.031	1.40891	1.40238	1.39942	1.39554
4.072	1.43985	1.43792	1.43409	1.42919

N	n 3018.5	3076.0	3082.0	3133.0
		18°		
0	1.35890	1.35740	1.35740	1.35622
0.5108	1.36803	1.36638	1.36637	1.36509
1.020	1.37691	1.37516	1.37510	1.37376
2.031	1.39432	1.39230	1.39220	1.39070
4.072	1.42765	1.42504	1.42494	1.42312

N	n 3255.0	3303.0	3345.6	3403.6
		18°		
0	1.35352	1.35200	1.35170	1.35062
0.5108	1.36216	1.36056	1.36021	1.35906
1.020	1.37066	1.36898	1.36851	1.36732
2.031	1.38711	1.38531	1.38475	1.38340
4.072	1.41868	1.41656	1.41588	1.41418

N	n 3467.0	3611.9	4415.9	4678.3
		18°		
0	1.34970	1.34754	1.33997	1.33831
0.5108	1.35805	1.35574	1.34770	1.34597
1.020	1.36625	1.36377	1.35527	1.35343
2.031	1.38214	1.37937	1.36999	1.36798
4.072	1.41260	1.40924	1.39802	1.39564

N	n 5086.0
18°	
0	1.33633
0.5108	1.34386
1.020	1.35116
2.031	1.36548
4.072	1.39264

Grufki, 1913

N	n H α	H β	H γ
	18°		
0	1.33142	1.33735	1.34058
0.5156	1.33889	1.34521	1.34868
1.020	1.34612	1.35269	1.35636
2.031	1.36008	1.36734	1.37147
4.072	1.38677	1.39529	1.40026

Heydweiller, 1912

N	κ	N	κ
	18°		
0.0998	117.2	1.010	1055
0.2000	226.1	2.009	2011
0.5006	540.9	4.000	3596

Clausen, 1914

t	λ 0.508 N	1.020 N	2.031 N	4.072 N
6	82.83	81.26	80.44	73.91
18	106.9	103.6	99.95	89.04
30	132.3	126.8	120.0	104.3

Heat constants.

Jauch, 1921

N	U	N	U
	18°		
0.5	0.9132	1.5	0.7642
1	0.8342	2	0.7006

H_2O + Rubidium iodide (RbI)

Heterogeneous equilibria .

Lannung, 1934

m	p	m	p
	18°		
0.02308	15.37	0.0928	15.04
0.02737	15.35	0.1739	14.67
0.03362	15.31	0.2702	14.21
0.04314	15.26	0.5574	12.94
0.04998	15.23	sat.d.	12.24
0.06627	15.16		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.921	1.4	0.878
0.2	0.904	1.6	0.880
0.3	0.896	1.8	0.882
0.4	0.890=	2.0	0.886
0.5	0.886	2.5	0.893
0.6	0.884	3.0	0.901
0.7	0.881	3.5	0.911
0.8	0.880	4.0	0.921
0.9	0.879	4.5	0.931
1.0	0.878	5.0	0.940
1.2	0.878		

Briggs, Conrad and Gregg, 1941

%	f.t.	%	f.t.
9.58	- 1.3	55.5	0
20.13	- 3.5	57.05	3.2
31.33	- 6.2	60.20	13.9
38.43	- 8.4	62.90	25.6
39.15	- 8.6	64.70	35.6
48.11	-12.2	66.76	48.5
49.79	-13.0	68.61	59.4
50.11	-13.0	70.89	77.2
51.18	-10.0	73.01	93.0
54.51	- 2.7	75.05	111.5

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.2692	0.2727	2.325	2.414
0.5738	0.5859	2.337	2.424
0.8313	0.8489	2.374	2.463
1.125	1.152	2.638	2.739
1.202	1.236	2.965	3.094
1.330	1.369	3.420	3.574
1.372	1.413	3.733	3.937
1.387	1.429	4.053	4.288
1.952	2.016	4.276	4.514
2.307	2.392	4.81	5.102

m_1 = m of KCl m_2 = m of RbI

Properties of phases

Getman, 1908

N	d	N	d
	25°		
0.264	1.0371	1.583	1.2475
0.528	1.0592	1.758	1.2662
0.792	1.1225	2.639	1.4149
1.056	1.1738	2.931	1.4612

Rubien, 1911 and Heydweiller, 1912

N	d	N	d
	18°		
0	0.99862	0.979	1.15664
0.0994	1.01481	2.017	1.32258
0.2012	1.01326	4.007	1.63622
0.5047	1.08033		

Buchanan, 1912

m	d	m	d
	19.5°		
0.5	1.0753	5	1.6055
1	1.1469	6	1.6944
2	1.2814	7	1.7876
3	1.4003	8.2307	1.18496(24.3)
4	1.5077		

N	d	N	d
	19.5°		
0.5	1.076665	0.008	0.999607
0.25	1.038085	0.004	0.998923
0.125	1.018349	0.002	0.998644
0.062	1.008402	0.001	0.998518
0.031	1.003404	0.000	0.997931
0.016	1.000873		

N	d	N	d
	23.0°		
0.5	-	0.008	0.998878
0.25	-	0.004	0.998265
0.125	1.017641	0.002	-
0.062	1.007682	0.001	-
0.031	1.002645	0.000	-
0.016	1.000163		

Grufki, 1913

N	d	N	d
	18°		
0	0.99862	2.039	1.32604
0.5115	1.08141	4.028	1.63939
0.031	1.16500		

Clausen, 1914

t	d			
	0.510 N*	1.025 N	2.025 N	4.015 N
6	1.08344	1.16748	1.33009	1.64778
18	1.08129	1.16429	1.32494	1.63996
30	1.07788	1.15991	1.32082	1.63318

*N at 18°

Baxter and Wallace, 1916

%	d	%	d
50.04°		25°	
63.45	1.86439	40.99	1.43992
40.31	1.42760	20.12	1.7574
20.29	1.16496	10.83	1.08620
10.92	1.07627	5.62	1.04150
5.67	1.03201		

%	d	%	d
	0°		
40.73	1.85009	10.79	1.09124
20.01	1.18259	5.60	1.04554

Getman, 1908

N	η	N	η
	25°		
0.264	871.4	1.583	788.7
0.528	848.9	1.758	791.8
0.792	830.3	2.639	779.6
1.056	814.7	2.931	787.0

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.35538
0.2	1.33771	2.0	1.37721
0.5	1.34450	4.0	1.41904

Dn $(H\gamma-D).10^3$ = 177

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.979	1.35492
0.0994	1.33542	2.017	1.37758
0.5047	1.34462	4.007	1.41921

Grufki, 1913

N	n H α	H β	H γ
	18°		
0	1.33742	1.33734	1.34053
0.5115	1.34262	1.34930	1.35301
0.031	1.35380	1.36131	1.36556
2.031	1.37536	1.38431	-
4.028	1.41627	1.42801	-

Clausen, 1914

t	λ 0.510 N	1.025 N	2.025 N	4.015 N
6	83.80	82.17	81.92	74.65
18	108.2	105.3	101.6	89.87
30	133.8	128.9	121.9	105.2

N at 18°

Heydweiller, 1912

N	κ	N	κ
	18°		
0.0994	116.6	0.9795	1038
0.2012	228.1	2.017	2055
0.5047	549.2	4.007	3615

H_2O + Rubidium chlorate ($RbClO_3$)

Calzolari, 1912

%	f.t.	%	f.t.
2.10	0	11.09	42.2
2.97	8	13.77	50
5.08	19.8	25.44	76
7.41	30	38.57	99

H_2O + Rubidium nitrate ($RbNO_3$)

Heterogeneous equilibria .

Tammann, 1885

%	p	%	p
		100°	
9.78	744.0	31.82	699.8
16.90	730.2	46.05	658.9
24.42	714.2		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
		25°	
0.1	0.903	1.2	0.725
0.2	0.871	1.4	0.706
0.3	0.847	1.6	0.689
0.4	0.826	1.8	0.673
0.5	0.809	2.0	0.656
0.6	0.794	2.5	0.620
0.7	0.781	3.0	0.588
0.8	0.768	3.5	0.561
0.9	0.756	4.0	0.538
1.0	0.745	4.5	0.516

Berkeley, 1904

%	f.t.	%	f.t.
16.94	0.60	68.70	63.40
30.68	15.85	73.96	75.60
46.42	31.55	79.29	90.95
58.24	45.85	86.05	118.3 b.t.

Jones, 1908

%	f.t.	%	f.t.
1.15	-	9.04	-3.5
1.23	-1.8	12.26	-4.2
5.11	-2.1		

%	spontaneous crystallization	%	spontaneous crystallization
22.94	7.0	54.29	39.3
28.65	12.6	57.20	41.5
32.20	16.2	59.29	46.7
32.98	16.8	62.44	50.8
33.17	17.3	64.75	55.1
38.97	22.5	66.64	57.5
39.32	24.0	68.00	61.1
41.51	25.6	72.84	70.0
43.42	27.8	73.07	66.9
43.33	28.0	75.45	76.6
46.45	30.9	77.52	82.4
48.35	30.3	79.44	86.1
51.16	36.1		

Properties of phases

Berkeley, 1904

%	t	d (sat.sol.)
16.94	0.60	1.1389
30.68	15.85	1.2665
46.42	31.55	1.4483
58.24	45.85	1.6216
68.70	63.40	1.8006
73.96	75.60	1.9055
79.29	90.95	2.0178
86.05	118.3 b.t.	2.1867

Buchanan, 1912

m	d	m	d
	19.5°		
0.5	1.0488	2	1.1842
1	1.0964	3	1.2654

N	d	N	d
	19.5°		
0.25	1.023143	0.016	0.99341
0.125	1.010648	0.008	0.998567
0. 0.062	1.004182	0.004	0.998017
0.031	1.000960	0	0.997931

Heydweiller, 1912 and Rubien, 1911

N	d	N	d
	18°		
0.1009	1.00922	0.9985	1.10115
0.2008	1.01964	1.984	1.19976
0.4996	1.05038	1.653	1.26584

Grufki, 1913

N	d	N	d
	18°		
0	0.99862	2.000	1.20134
0.5044	1.05091	2.691	1.26950
1.006	1.10190		

Clausen, 1914

t	d			
	0.5035 N	1.008 N	2.000 N	2.685 N
6	1.05339	1.10558	1.20652	-
18	1.05087	1.10226	1.20161	1.26924
30	1.04719	1.09783	1.19592	1.26260

N at 18°

Smith, Wolfenden and Hartley, 1931

b	d	
	18°	25°
0.09244	1.00838	1.00671
0.14440	.01490	.01318
0.18531	.01832	.01656
0.25122	.02548	.02363
0.36593	.03811	.03611
0.47418	.05025	.04816
0.47581	.05058	.04848
0.71402	.07818	.07581
0.89168	.09962	.09704
0.98708	.11135	.10868
1.26927	.14461	.14461
1.5505	.18620	.18287
1.78637	.22030	.21668

b = mol - gr in 1000 g solution

Pesce, 1934

N	λ	N	λ
		18°	
0.189	27.50	0.981	41.48
0.258	31.80	1.097	42.05
0.498	37.90	1.4577	43.18
0.499	37.64	1.839	43.90
0.770	40.32	2.180	44.46
		25°	
0.188	28.25	0.978	42.12
0.257	32.05	1.094	42.66
0.497	38.62	1.452	43.75
0.490	38.39	1.834	44.42
0.768	41.00	2.174	44.93

Smith, Wolfenden and Hartley, 1931

b	η ($water^t$ = 1)	
	18.00°	25.01°
0.09244	0.99251	0.99422
0.14440	.98770	.99100
0.18531	.98386	.98850
0.25122	.97846	.98490
0.47418	.96278	.97383
0.47581	.96225	.97326
0.71402	.94826	.96308
0.89168	.94100	.95877
0.98708	.93740	.95629
1.26927	.93128	.95421
1.5505	.93060	.95715
1.78637	.93388	.96394

b = mol - gr in 1000 g solution,.

Heydweiller, 1909

N	n_D	N	n_D
		18°	
0	1.33327	1.0	1.34360
0.1	1.33435	2.0	1.35308
0.2	1.33542	4.0	1.36216
0.5	1.33855		

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.9985	1.34359
0.1009	1.33438	1.989	1.35296
0.2008	1.33542	2.653	1.35900
0.4996	1.33854		

Grufki, 1913

N	n H_α	n H_β	n H_γ
0	1.33139	1.33733	1.34048
0.5044	1.33674	1.34288	1.34620
1.006	1.34176	1.34811	1.35158
2.000	1.35110	1.35783	1.36153
2.691	1.35728	1.36425	1.36814

Heydweiller, 1912

N	$\varkappa$	N	$\varkappa$
	18°		
0.1009	113.6	0.9985	855.4
0.2008	212.4	1.984	1467
0.4996	475.6	2.653	1791

Clausen, 1914

t	λ 0.5035 N	λ 1.008 N	λ 2.000 N	λ 2.685 N
6	72.63	65.45	57.12	-
18	95.17	85.23	73.48	66.96
30	118.8	105.8	90.43	82.42

N at 18°

H_2O + Rubidium perchlorate ($RbClO_4$)

Calzolari, 1912

%	f.t.	%	f.t.
0.58	8.5	4.62	60
0.76	14	6.23	70
1.80	33.7	9.43	84
2.57	42	14.81	99
3.43	50		

H_2O + Rubidium sulfate (Rb_2O_4S)

Tammann, 1885

%	p	%	p
		100°	
10.37	747.3	29.52	715.9
16.00	739.4	39.30	689.5
23.93	726.9		

Etard, 1894 (fig.)

%	f.t.	%	f.t.
27.4	3	43.9	97
32.5	20	49.2	170
37.3	37		

Berkeley, 1904

%	f.t.	%	f.t.
26.83	0.50	39.97	57.90
31.53	15.80	42.28	74.75
35.17	31.60	44.01	89.45
37.79	44.20	45.22	102.4 b.t.

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
		25°	
0.1	0.799	0.8	0.691
0.2	0.764	0.9	0.686
0.3	0.740	1.0	0.681
0.4	0.724	1.2	0.677
0.5	0.714	1.4	0.677
0.6	0.705	1.6	0.679
0.7	0.698	1.8	0.684

Berkeley and Applebey , 1911

sat.sol. b.t. = 103.511°

Rubien, 1911 and Heydweiller, 1912

N	d	N	d
	18°		
0	0.99862	0.9955	1.10601
0.1003	1.00991	2.003	1.21032
0.2009	1.02093	3.135	1.32418
0.4953	1.05294		

Grufki, 1913

N	d	N	d
	18°		
0	0.99862	2.044	1.21437
0.5029	1.05376	3.166	1.32712
1.009	1.10744		

Clausen, 1914

t	d			
	0.501 N*	1.010 N	2.043 N	3.168 N
6	1.05584	1.17044	1.21885	1.33273
18	1.05357	1.10756	1.21472	1.32769
30	1.04996	1.10371	1.20999	1.32301

* N at 18°

Tunha, 1914 - 1915

N	d				
	10°	20°	30°	40°	50°
0	0.99973	0.99823	0.99567	0.99224	0.98807
0.1	1.0103	1.0086	1.0068	1.0029	0.9987
0.2	.0213	.0185	.0166	.0135	1.0080
0.5	.0535	.0512	.0486	.0449	.0406
0.8	.0845	.0822	.0792	.0757	.0714
1.0	.1049	.1020	.0985	.0948	.0905
1.5	.0546	.1516	.1476	.1435	.1390
2.0	.2027	.1987	.0945	.1898	.1852

N	η ($water^t = 1$)				
	10°	20°	30°	40°	50°
0.1	0.9976	0.9849	1.0089	1.0148	1.0141
0.2	0.9999	.9908	1.0169	.0195	.0304
0.5	1.0103	1.0099	.0462	.0582	.0680
0.8	1.0281	.0342	.0800	.0858	.1071
1.0	1.0354	.0509	.0996	.1066	.1333
1.5	1.0930	.0939	.1519	.1634	.1968
2.0	1.1216	.1394	.2093	.2297	.2654

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34472
0.1	1.33453	2.0	1.35482
0.2	1.33575	3.0	1.36377
0.5	1.33925		

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.9955	1.34467
0.1003	1.334535	2.003	1.35485
0.2009	1.33576	3.135	1.36512
0.4953	1.33919		

Grufki, 1913

N	n H α	H β	H γ
		18°	
0	1.33139	1.33733	1.34054
0.5029	1.33737	1.34340	1.34664
1.009	1.34289	1.34894	1.35221
2.044	1.35324	1.35943	1.36273
3.166	1.36342	1.36970	1.37306

Clausen, 1914

t	0.501 N*	λ 1.010 N	2.043 N	3.168 N
6	65.00	59.54	54.29	48.92
18	85.38	77.53	69.52	62.03
30	106.9	96.34	85.37	75.60

* N at 18°

Heydweiller, 1912

N	ϰ	N	ϰ
	18°		
0.1003	104.1	0.9955	780.2
0.2009	192.8	2.005	1403
0.4953	424.9	3.138	1960

H_2O + Rubidium acid sulfate ($RbHO_4S$)

Tammann, 1885

%	p	%	p
		100°	
8.37	749.1	30.06	709.0
16.58	736.1	48.82	651.5
23.14	723.9		

H_2O + Rubidium selenate (Rb_2SeO_4)

Tutton, 1897

%	d
20°	
40.60	1.4688
47.07	1.5806

61.39% f.t. =12°

H_2O + Rubidium pentaborate (RbB_5O_8)

Rollet and Andres, 1931

%	f.t.	%	f.t.
1.57 E	-0.43	5.75	45
1.58	+0.2	8.69	60
1.72	5	11.5	70.4
2.21	13.4	15.2	82.4
2.57	18	17.75	88.15
3.58	30	23.75	102

sat.sol b.t. = 102.17°

H_2O + Rubidium telluroxalate ($Rb_2C_2H_6O_{10}Te$)

Rosenheim and Weinheber, 1911

%	f.t.	%	f.t.
3.71	0.	11.31	40.
7.20	20.	14.46	50.
8.59	30.		

H_2O + Rubidium formate ($RbCHO_2$)

Sidwick and Gentle, 1922

%	f.t.	%	f.t.
6.23	- 1.82	84.61	16.3
15.93	- 5.62	-	16.5 tr.t.
20.62	- 7.82	85.60	28.3 (1+2)
25.79	-10.62	87.77	43.6
30.62	-13.88	89.23	49.9
78.86	+ 3.3 (1+1)	-	51.0 tr.t.
80.71	7.8	90.06	60.8 0aq
81.37	9.5	101.7	101.7
83.59	14.0	170	170

H_2O + Rubidium acetate ($RbC_2H_3O_2$)

Sidwick and Gentle, 1922

%	f.t.	%	f.t.
9.59	- 2.79	86.25	+ 44.7
16.11	- 5.27	89.30	99.4
19.64	- 6.82	91.35	125.2
25.57	-10.27	100	246
82.90	- 9.50		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.943	1.0	1.023
0.2	0.945	1.2	1.046
0.3	0.952	1.4	1.068
0.4	0.961	1.6	1.091
0.5	0.971	1.8	1.114
0.6	0.981	2.0	1.137
0.7	0.992	2.5	1.192
0.8	1.002	3.0	1.248
0.9	1.013	3.5	1.302

H_2O + Rubidium tartrate ($Rb_2C_4H_4O_4$)

Rimbach, 1895 and Pribram and Glücksmann, 1897

%	d	%	d
	20°		
64.49	1.7379	20.26	1.1635
59.56	1.6552	10.25	1.0763
54.05	1.5696	5.11	1.0363
49.43	1.4990	1.57	1.0098
29.84	1.2583	0	0.9982
29.84	1.2583		

Rimbach, 1895

%	(α)mol red	D	green	pale blue	dark blue	mol. cond.
			20°			
64.49	18.06	23.47	29.03	35.06	41.43	-
59.56	17.94	23.17	28.77	34.69	41.00	-
54.05	17.74	22.85	28.42	34.18	40.61	-
49.43	17.42	22.47	27.90	33.71	39.97	-
40.03	17.06	21.91	27.27	33.00	38.99	-
29.84	16.59	21.35	26.55	31.95	37.67	44.93
20.26	16.13	20.73	25.95	31.35	36.95	53.54
10.25	15.60	20.15	25.00	30.35	35.99	63.93
5.11	-	19.98	-	-	-	71.24
1.57	-	19.54	-	-	-	83.18

Pribram and Glücksmann, 1897

%	d	$(\alpha)_D$
	20°	
10.2510	1.07663	20.19
8.5655	1.06338	20.07
6.1771	1.04477	19.86
5.1379	1.03662	19.77
3.2380	1.02230	19.36
1.5296	1.00951	19.01
0.5266	1.00202	18.76
0	0.99823	-

H_2O + Rubidium saccharate ($RbC_6H_{11}O_6$)

Rimbach and Heiten, 1908

c	(α)mol red	yellow	green	pale blue	dark blue
8.00	-2.02	-3.71	-4.34	-5.58	- 7.45
12.00	-2.27	-3.96	-4.70	-6.07	- 8.28
16.00	-2.53	-4.25	-5.37	-6.78	- 9.12
20.00	-2.82	-4.65	-6.15	-7.92	-10.30
24.00	-3.58	-5.25	-6.82	-8.40	-11.13

H_2O + Rubidium benzoate ($RbC_7H_5O_2$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
9.41	- 1.56	56.06	+15.0
22.15	- 4.71	59.70	51.5
33.32	- 9.04	63.23	82.0
45.75	-14.75	70.32	147.0

H_2O + Rubidium salicylate ($RbC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
12.34	- 1.62	68.15	21.5
23.48	- 3.69	69.84	27.0
33.81	- 6.34	74.97	49.0 0 aq.
42.77	- 8.82	77.96	73.0
53.40	-13.17	79.89	88.0
64.85	+ 8,0 (1+1)	86.36	134.0

H_2O + Rubidium m-hydroxybenzoate ($RbC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
14.92	- 2.23	60.02	45.0
32.13	- 6.41	65.04	64.0
48.88	-13.69 E	70.45	88.0
50.61	+14.0 (1+1)	74.49	105.0
54.94	32.5 0aq.	79.91	130.0

H_2O + Rubidium p-hydroxybenzoate ($RbC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
9.92	- 1.17	55.73	68.0
17.47	- 2.47	64.95	94.0 0 aq.
26.22	- 4.22	74.00	120.5
35.66	+17.5	75.92	127.0
45.95	45.0		

H_2O + Cesium fluoride (CsF)

Lannung, 1934

m	p	m	p
	18°		
0.773	15.08	20.85	2.77
1.103	14.94	20.90	2.74
1.784	14.60	22.94	1.89
1.813	14.44	27.92	1.34
3.651	13.17	28.67	0.96
4.325	12.84	31.65	0.69
6.968	10.73	36.47	0.59
11.01	7.29	37.25	0.55
12.67	6.45	satd	0.35
13.36	5.72		

p dissoc. (1+1) = 0.002

H_2O + Cesium chloride (CsCl)

Heterogeneous equilibria

Tammann, 1885

t	p 0%	p 22.43%
30.20	32.2	30.3
44.92	71.6	66.8
48.20	84.6	78.1
51.53	99.8	93.1
53.97	112.4	105.0
56.52	127.0	118.4
61.53	160.4	150.4
65.11	188.6	176.9
67.76	212.1	199.2
71.06	244.7	230.0
74.67	285.3	269.4
78.66	336.7	317.4
81.86	382.5	360.2
83.87	415.0	390.8
88.06	488.4	460.0
90.07	527.3	496.8
93.37	599.2	564.9
95.77	652.1	613.2
100.00	760.1	719.0

Lannung, 1934

m	p	m	p
	18°		
0.6503	15.18	2.781	14.19
0.705	15.17	4.241	13.48
0.908	15.06	6.114	12.58
1.117	14.98	8.77	11.35
1.290	14.89	10.72	10.48
1.647	14.72	sat.d.	10.43
2.044	14.54		

Morey and Chen, 1956

t	P
V + L + C	
400	65
500	65
600	29

Foote, Saxton and Dixon, 1932

Sat. sol. $\lg p = 2198.5 / T + 8.5621$

Buchanan, 1899

%	b.t.	%	b.t.
	746.55 mm		
46.97	105.06	14.46	100.41
40.23	103.53	12.97	100.31
34.75	102.52	11.64	100.21
27.04	101.51	9.73	100.11
22.99	101.10	8.00	100.01
20.70	100.90	6.54	99.91
18.45	100.69	0.00	99.51
15.75	100.51		

Berkeley and Applebey, 1911

Sat. sol. b.t. = 119.919°

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.917	1.4	0.856
0.2	0.897	1.6	0.857
0.3	0.885	1.8	0.859
0.4	0.875	2.0	0.862
0.5	0.869	2.5	0.869
0.6	0.864	3.0	0.879
0.7	0.861	3.5	0.889
0.8	0.859	4.0	0.900
0.9	0.858	4.5	0.912
1.0	0.857	5.0	0.924
1.2	0.856		

Berkeley, 1904

%	f.t.	%	f.t.
61.86	0.70	69.65	60.20
64.56	16.20	71.07	76.10
66.34	29.85	72.18	89.50
68.09	45.55	74.36	119.4

Hinrichsen and Sachsel, 1904

%	f.t.	%	f.t.
61.9	0.3	66.3	30
63.5	10	67.4	40
64.9	20		

Mariani and Di-Giacomo-Dejak, 1950

N	f.t.	N	f.t.
0.064178	-0.21873	0.224924	-0.7402
0.061198	-0.20810	0.853062	-2.707
0.095162	-0.32207	0.982053	-3.141
0.116592	-0.3966	1.179012	-3.753
0.147435	-0.4949	1.568669	-5.048
0.197305	-0.6547	1.895922	-6.235

Lyons and Riley, 1954

N	D	N	D
	25°		
0.0625	1.887	1.960	2.023
.0900	1.874	3.060	2.183
.1600	1.859	4.000	2.291
.2500	1.855	5.000	2.364
.3600	1.855	5.750	2.354
.6400	1.868	6.000	2.335
1.000	1.902		

Properties of phases

Berkeley, 1904

%	t	d (sat.sol.)
61.86	0.70	1.8458
64.56	16.20	1.8984
66.34	29.85	1.9359
68.09	45.55	1.9702
69.65	60.20	2.0012
71.07	76.10	2.0286
72.18	89.50	2.0500
74.36	119.4	2.0859

Heydweiller, 1909

N	d	N	d
	18°		
0.1	1.01151	1.0	1.1265
0.2	1.02450	2.0	1.2518
0.5	1.06293		

Schneider, 1910

N	d	N	d
	18°		
2.0	1.2523	0.2	1.02454
1.0	1.1266	0.1	1.01153
0.5	1.06305	0	0.99862

Buchanan, 1912

m	d	m	d
	19.5°		
0.5	1.0599	6	1.5644
1	1.1060	7	1.6447
2	1.2250	8	1.6997
3	1.3223	9	1.7561
4	1.4090	12.1563	1.9055 (23°)
5	1.4931		

N	d	N	d
	19.5°		
0.5	1.060842	0.016	1.000396
0.25	1.030059	0.008	0.999395
0.125	1.014340	0.004	.998885
0.062	1.006395	0.002	.998620
0.031	1.002400	0.000	.997931

Grufki, 1913 and Lübben, 1913

N	d	N	d
	18°		
0	0.99862	2.001	1.25200
0.5026	1.06326	3.994	1.49438
0.999	1.12631		

Clausen, 1914

t	d			
	0.504 N*	1.002 N	2.007 N	3.994 N
6	1.06553	1.12959	1.25702	1.50511
18	1.06344	1.12685	1.25312	1.49955
30	1.06014	1.12310	1.24861	1.49403

N at 18°

Baxter and Wallace, 1916

%	d	%	d
70.19°		50.04°	
66.41	1.90437	66.14	1.91401
47.81	1.51306	47.54	1.22215
28.92	1.24737	28.71	1.75732
13.30	1.08665	13.20	1.09685
6.96	1.03214	6.88	1.04240
3.56	1.00487	3.53	1.01516
25°		0°	
65.80	1.92554	46.96	1.54367
47.22	1.53400	28.36	1.27523
28.50	1.26769	13.02	1.11177
13.08	1.10654	6.81	1.05581
6.82	1.05174	3.49	1.02784
3.50	1.02437		

Geffcken, 1929

m	d	m	d
	25°		
0	0.99707	7.8454	1.71708
1.9592	1.22520	10.1726	1.85829
5.9584	1.58372		

Rabinovitsch, 1921

%	d	%	d
	0°		
62.13	1.8573	32.60	1.3278
59.30	.7850	18.70	.1667
54.36	.6800	9.88	.0224
48.64	.5711	4.50	.0360
40.71	.4404	2.20	.0172

%	d	%	d
	50°		
67.63	1.9492	31.30	1.2735
61.94	.8062	25.35	.2080
54.96	.6569	19.53	.1902
48.17	.5329	14.62	.1044
43.22	.4930	9.81	.0645
38.45	.3831		

%	d	%	d
	100°		
72.5	2.037	62.3	1.767
67.2	1.893	55.5	1.626

Lyons and Riley, 1954

N	d	N	d
	25°		
0.5002	1.0610	3.998	1.4966
1.000	1.1243	5.001	1.6197
2.269	1.2829	5.998	1.7408
2.997	1.3733		

Pohl, 1852

%	relative	%	π relative
	13.9°		
0	1.000	33.50	748
13.20	0.893	38.46	709
23.20	830		

Schneider, 1910

N	η (water=1)	N	η (water = 1)
	18°		
2.0	0.9230	0.2	0.9883
1.0	0.9510	0.1	0.9940
0.5	0.9831		

Rabinovitsch, 1921

%	η ($water^t$=1)	%	η ($water^t$=1)
	50°		
67.63	1.779	31.30	1.069
61.94	1.504	25.35	.042
54.96	1.306	19.53	.023
48.17	1.193	14.62	..015
43.22	1.140	9.81	.012
38.45	1.107		

%	η ($water^t$=1)	%	η ($water^t$= 1)
	0°		
62.13	0.909	32.60	0.783
59.30	.856	18.70	.849
54.36	.799	9.88	.910
48.64	.766	4.50	.954
40.71	.761	2.20	.974

%	η ($water^t$=1)	%	η ($water^t$=1)
	100°		
72.5	2.691	62.3	1.855
67.2	2.540	55.5	1.602

Lyons and Riley, 1954

N	η	N	η
	25°		
0.5002	875.1	3.998	878.0
1.000	861.0	5.001	926.3
2.269	846.1	5.998	1009
2.997	851.5		

Rabinovitsch, 1921

%	κ	%	κ
	0°		
62.13	3301	32.60	1714
59.30	3273	18.70	872
54.36	3111	9.88	431
48.64	2800	4.50	193
40.71	2268	2.20	96

%	κ	%	κ
	50°		
67.63	5850	31.30	3340
61.94	6050	25.35	2660
54.96	5840	19.53	2020
48.17	4320	14.62	1510
43.22	4800	9.81	1020
38.45	4230		

%	κ	%	κ
	100°		
72.5	7720	62.3	8570
67.2	8360	55.5	8360

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	2.0	1.35853
0.5	1.33986	4.0	1.38187
1.0	1.34623		

Grufki, 1913

N	n Hα	n Hβ	n Hγ
		18°	
0	1.33141	1.33736	1.34054
0.5026	1.33798	1.34412	1.34742
0.999	1.34426	1.35055	1.35399
2.001	1.35648	1.36316	1.36884
3.994	1.37953	1.38690	1.39105

Lubben, 1913

N	n			
	2573.2 A°	2748.7	2881.1	2981.1
		18°		
0	1.37361	1.36654	1.36260	1.35980
0.5026	1.38176	1.37432	1.37017	1.36725
0.999	1.38962	1.38180	1.37747	1.37438
2.001	1.40516	1.39661	1.39186	1.38855
3.994	1.43414	1.42421	1.41869	1.41488

N	n			
	3133.0 Å	3467.0	3611.9	5086.0
		18°		
0	1.35622	1.34970	1.34734	1.33633
0.5026	1.36353	1.35674	1.35447	1.34283
0.999	1.37050	1.36346	1.36113	1.34899
2.001	1.38435	1.37675	1.37423	1.36125
3.994	1.41007	1.40149	1.39861	1.38378

Geffcken, 1929

m	n_D	m	n_D
	25°		
0	1.33253	7.8454	1.40093
1.9592	1.35508	10.1726	1.41367
5.9584	1.38869		

Heydweiller, 1909

N	λ	N	λ
	18°		
0.1	115.0	1.0	100.2
0.2	110.0	2.0	95.8
0.5	104.1		

Clausen, 1914

t	λ			
	0.504 N*	1.002 N	2.007 N	3.994 N
6	80.7	80.5	77.1	71.5
18	104.1	99.8	95.5	85.5
30	129.5	122.7	115.3	100.8

*N at 18°

Okazaki, 1942

%	Verdet's constant (D)	%	Verdet's constant (D)
	20°		
11.15	0.01389	34.35	0.01611
16.04	.01429	40.23	.01684
21.94	.01483	46.01	.01758
27.60	.01537		

Okazaki, 1933

%	Verdet's constant (3555 Å)
	28°
8,14	0.04290
16.45	.04496
22.41	.04668
31.31	.05011
39.25	.05298

Kapustinskii and Ruzavin, 1955

Thermic conductivity coefficient

%	$k.10^6$
	25°
6.35	1422
15.4	1376
20.5	1350

H_2O + Cesium bromide (CsBr)

Lannung, 1934

m	p	m	p
	18°		
0.3834	15.31	1.069	15.00
0.4142	15.30	1.476	14.81
0.4511	15.26	1.802	14.68
0.5230	15.24	2.885	14.22
0.6208	15.20	3.696	13.89
0.7705	15.13	4.664	13.47
0.926	15.57	satd	13.24

Buchanan, 1912

%	f.t.	%	f.t.
48.19	7.5	60.93	50
54.75	24.5	68.95	93.5

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.2021	0.2060	2.841	3.057
.2945	.3011	3.054	3.284
.5978	.6240	3.244	3.504
.8222	.8614	3.605	3.888
1.184	1.257	3.974	4.287
1.554	1.658	4.061	4.366
1.692	1.805	4.214	4.535
1.954	2.090	4.469	4.799
2.302	2.470	4.768	5.104

m_1 = molality of potassium chloride

m_2 = " " cesium bromide

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.917	1.4	0.848
0.2	0.896	1.6	0.848
0.3	0.882	1.8	0.850
0.4	0.873	2.0	0.852
0.5	0.865	2.5	0.859
0.6	0.861	3.0	0.866
0.7	0.857	3.5	0.874
0.8	0.854	4.0	0.884
0.9	0.852	4.5	0.892
1.0	0.850	5.0	0.901
1.2	0.849		

Buchanan, 1912

N	d	
	21.4°	23.0°
0.5	1.079175	-
0.25	1.039316	-
0.125	1.019040	1.018237
0.062	1.008764	1.007975
0.031	1.003545	1.002847
0.016	1.000999	1.000242
0.008	1.999640	0.998943
0.004	0.998978	-
0.002	0.998679	-
0.001	0.998517	-
0.000	0.997931	0.997565

m	d	m	d
	21.4°		
1	1.1467	4	.5516
2	.3015	5	.6590
3	.4297	5.3057	.6968

Baxter and Wallace, 1916

%	d	%	d
70.19°		50.04°	
53.70	1.65596	53.39	1.66651
47.18	1.47667	43.90	1.48720
26.47	1.22787	26.28	1.23825
12.10	1.07839	11.99	1.08872
6.31	1.02788	6.35	1.03827
3.22	1.00275	3.19	1.01313
25 °		0°	
53.09	1.67853	43.34	1.50818
43.53	1.49871	25.95	1.25584
26.08	1.24875	11.86	1.10327
11.94	1.09841	6.18	1.05151
6.20	1.04765	3.14	1.02564
3.15	1.02230		

Heydweiller, 1921

N	d 18°	λ
0.5	1.0836	106.4
1	1665	101.8
2	3314	97
3	4948	89.9
4	6571	84.2

Gibson, 1935

%	π (1-1000 bars)	d
	25°	
0	39.35	0.99701
9.88	37.70	1.07992
20.05	36.00	1.1804
29.94	34.18	1.29012
-	34.05	-
40.00	32.05	1.43953
-	31.98	-
45.69	30.81	1.5354
49.77	29.82	1.6120

Jauch, 1921

N	U	N	U
		18°	
0.5	0.8881	3	0.5290
1	0.7980	4	0.4608
2	0.6302		

H_2O + Cesium iodide (CsI)

Lannung, 1934

m	p	m	p
		18°	
0.3781	15.29	1.264	14.91
0.4146	15.29	1.431	14.85
0.5223	15.23	1.804	14.67
0.5588	15.19	2.068	14.57
0.755	15.15	2.447	14.41
0.919	15.11	sat.d.	14.31
1.069	15.04		

Foote, 1903

%	f.t.
27.69 E	- 4.0
51.48	+35.6

Robinson, 1935

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
		25°	
0.2053	0.2090	1.618	1.886
.3777	.3910	1.662	1.943
.5145	.5376	1.706	2.000
.6526	.6954	1.762	2.049
1.000	1.096	1.860	2.187
1.181	1.322	2.092	2.473
1.411	1.609	2.315	2.738
1.584	1.838	2.552	3.032
1.598	1.862		

m^1 = molality of potassium chloride

m 2 = " " cesium iodide

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
		25°	
0.1	0.916	1.0	0.846
0.2	0.895	1.2	0.842
0.3	0.880	1.4	0.839
0.4	0.870	1.6	0.836
0.5	0.863	1.8	0.834
0.6	0.858	2.0	0.832
0.7	0.855	2.5	0.827
0.8	0.852	3.0	0.822
0.9	0.849		

Buchanan, 1912

%	f.t.
33.77	12
61.08	107

Briggs and Hubard, 1941

%	f.t.
8.76	- 1.2
18.75	- 2.7
27.45 E	- 4.0

%	f.t.	%	f.t.
31.41	1.4	60.43	59.3
36.90	9.3	60.75	61.3
43.32	19.4	65.24	77.7
49.98	32.0	67.16	88.0
50.05	32.4	70.25	102.8
55.54	45.9	71.48	109.1

Buchanan, 1912

m	d	m	d
	23.6°		
1	1.1718	3.385	1.5389 (23.1°
2	.3431	3.5454	.5487 (22.8°
3	.4887		

N	d 19.5°	d 23.0°
0.5	1.096000	-
0.25	.048131	-
0.125	.023304	1.022635
0.062	.010880	.010221
0.031	.004661	.003940
0.016	.001487	.000769
0.008	.999915	0.999206
0.004	.999108	.998526
0.002	.998644	-
0.001	.998472	-

Baxter and Wallace , 1916

%	d	%	d
70.19°		50.04°	
48.31	1.55084	48.15	1.56209
14.82	1.26431	14.70	1.27517
13.35	1.09277	13.62	1.10333
7.22	1.03507	7.16	1.04553
3.70	1.00622	3.65	1.01669
25°		0°	
47.78	1.57541	14.48	1.29411
14.58	1.28630	13.46	1.11850
13.50	1.11332	7.05	1.05906
7.08	1.05505	3.61	1.02936
3.62	1.02594		

Jones and Fornwalt, 1936

N	d	N	d
	0°		
0.000200	0.99992	0.078389	1.01605
.000500	1.00002	.099981	1.02052
.001000	1.00011	.150018	1.03077
.002000	1.00036	.200017	1.04168
.005000	1.00091	.2885023	1.05849
.009998	1.00195	.385037	1.07896
.010000	1.00193	.496900	1.10182
.019977	1.00401	.750096	1.15341
.029150	1.00607	1.00007	1.23415
.049993	1.01021	1.43997	1.29307
.074995	1.01538		

N	d	N	d
	25°		
0.000939	0.99710	0.099661	1.01717
.000998	0.99714	.149505	1.02724
.000997	.99726	.199281	1.03725
.001994	.99746	.283884	1.05425
.004985	.99806	.383349	1.07422
.009971	.99908	.494697	1.09656
.019919	1.00108	.745835	1.14689
.029862	1.00309	.993688	1.19645
.049988	1.00713	1.42912	1.28333
.074758	1.01216	1.99969	1.39696
.078141	1.01284		

N	η(water=1)	N	η (water=1)
	0°		
0.000200	1.00001	0.078389	0.98117
.000500	0.99998	.099981	.97000
.001000	.99985	.150018	.96441
.002000	.99962	.200017	.95329
.005000	.99851	.2885037	.93480
.009998	.99773	.3885023	.91467
.010000	.99771	.496900	.89346
.019977	.99535	.750096	.85166
.029150	.99289	1.00007	.81561
.049993	.98802	1.43997	.76524
.074995	.98205		

N	η(water=1)	N	η(water=1)
	25°		
0.000199	1.00009	0.099661	0.98910
.000498	0.99997	.149505	.98344
.000997	1.00003	.199281	.97785
.001994	0.99983	.283884	.96897
.004985	.99963	.383349	.95866
.009971	.99911	.494697	.94799
.019919	.99815	.745835	.92616
.029862	.99709	.993688	.90726
.049988	.99478	1.42912	.88080
.074758	.99192	1.99969	.85818
.078141	.99160		

Kapustinskii, Yakuchevskii and Drakin, 1953

%	U	%	U
	25°		
0	0.9980	12.90	0.8561
3.72	0.9567	15.98	0.8228
10.40	0.8834		

Kapustinskii, Lipilina and Samoilov, 1956

m	U	m	U
	25°		
1.0614	0.7655	0.4275	0.8883
0.8235	0.8058	0.2212	0.9393
0.6306	0.8456		

H_2O + Cesium hydroxide (CsOH)

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.942	0.6	0.970
0.2	0.939	0.7	0.982
0.3	0.940	0.8	0.992
0.4	0.949	0.9	1.003
0.5	0.960	1.0	1.014

H_2O + Cesium chlorate ($CsClO_3$)

Calzolari, 1912

%	f.t.	%	f.t.
2.40	0	13.00	42.2
3.38	8	16.25	50
5.91	19.8	29.40	77
8.70	30	43.34	99

H_2O + Cesium nitrate (CsO_3N)

Johnston, 1907

N	b.t. molar rise	N	b.t. molar rise
	99.4°		
0.312	0.99	1.965	0.94
0.758	0.98	2.202	0.94
1.175	0.86	2.378	1.02
1.418	0.92	2.502	1.05
1.725	0.92	2.680	1.07

Berkeley, 1904

%	f.t.	%	f.t.
8.71	0.35	45.50	59.90
16.29	15.95	55.48	76.40
25.65	30.45	62.29	90.55
35.75	45.15	68.78	106.2 b.t.

Berkeley and Applebey, 1911

Sat sol. b.t. = 107.228°

Jones, 1908

%	f.t.	%	f.t.
0.21	-1.2	5.67	-3.2
1.23	-2.5	7.41	-3.2
3.84	-3.0		

%	spontaneous crystallization	%	spontaneous crystallization
9.03	- 0.3	28.23	33.8
10.83	4.5	30.74	38.4
11.85	6.5	32.66	39.5
15.14	13.6	35.75	45.3
16.88	16.9	38.71	48.5
18.63	19.5	41.30	53.3
21.21	23.5	47.38	62.9
23.59	27.5	49.84	65.1
25.47	29.9	56.54	78.7

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.902	0.7	0.774
0.2	0.869	0.8	0.761
0.3	0.842	0.9	0.748
0.4	0.820	1.0	0.736
0.5	0.802	1.2	0.715
0.6	0.787	1.4	0.695

Berkeley, 1904

%	t	d
8.71	0.35	1.0701
16.29	15.95	1.1345
25.65	30.45	1.2219
35.75	45.15	1.3306
45.50	59.90	1.4565
55.48	76.40	1.6068
62.29	90.55	1.7307
68.78	106.2 b.t.	1.8657

Merton, 1910

%	d	%	d
		0°	
0	0.99987	3.8654	1.03023
0.9975	1.00757	5.8813	1.04654
1.9400	1.01491	7.7256	1.06195
		10°	
0	0.99973	5.4227	1.04191
1.0358	1.00755	8.1634	1.06445
2.1382	1.01598	10.7699	1.08687
3.1347	1.02375	12.5140	1.10207
		18°	
0	0.99862	6.1810	1.04633
2.0966	1.01437	8.8065	1.06805
3.0349	1.02157	11.0114	1.08671
4.8234	1.03555	14.2629	1.11525
		25°	
0	0.99707	6.1810	1.04436
0.9960	1.00443	8.8065	1.06580
2.0966	1.01268	11.0114	1.08441
3.0349	1.01981	14.2629	1.11298
4.8234	1.03369		

Heydweiller, 1921

N	d	N	d
	18°		
0.1	1.01454	0.5	1.0724
0.2	1.0290	1	1.1434

Johnston, 1907

N	λ	N	λ
	99.4°		
2	1725	0.5	2485
1.5	1990	0.10	2889
1	2206	0.001	3777

Merton, 1910

%	η	%	η
	0°		
(0.4896)	177.09	3.8654	170.63
0.9975	176.03	(5.5772)	167.86
(1.4826)	174.97	(5.8813)	164.36
	10°		
(0.5691)	129.86	5.4227	123.14
1.0358	129.36	8.1634	121.25
3.1347	127.46	10.7699	119.97
(2.1382)	125.36		
	18°		
(0.4509)	231.4	6.1810	3173
0.9960	511	8.8065	4520
2.0966	1076	11.0114	5652
3.0349	1557	14.2629	7321
4.8234	2475		
	25°		
0.4509	88.99	6.1810	86.97
0.9960	88.83	8.8065	86.17
2.0966	88.44	11.0114	85.57
3.0349	88.04	14.2629	84.80
4.8234	87.42		

Heydweiller, 1921

N	λ	N	λ
	18°		
0.1	106.5	0.5	91.1
0.2	100.2	1	82

Jauch, 1921

N	U
18°	
0.2	0.9584
0.5	0.9035
1	0.8420

H_2O + Cesium perchlorate ($CsClO_4$)

Calzolari, 1912

%	f.t.	%	f.t.
0.90	8.5	6.80	60
1.18	14	8.92	70
2.90	33.7	14.16	84
3.93	42	22.22	99
5.18	50		

H_2O + Cesium sulfate (Cs_2O_4S)

Morey and Chen, 1956

t	P kg	t	P kg
	V + L + C		
400	144	600	187
500	186	700	92

560° maximum P

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.804	0.8	0.717
0.2	0.772	0.9	0.714
0.3	0.751	1.0	0.712
0.4	0.739	1.2	0.711
0.5	0.731	1.4	0.713
0.6	0.725	1.6	0.716
0.7	0.721	1.8	0.722

Berkeley and Applebey, 1911

Sat. sol. b.t. = 109.964°

Berkeley, 1904

%	f.t.	d
	sat. sol.	
62.62	0.70	1.9766
62.96	15.00	1.9992
64.83	30.40	2.0202
65.82	44.90	2.0365
66.60	59.50	2.0512
67.53	75.70	2.0664
68.23	89.75	2.0774
69.18	108.6 b.t.	2.0932

Tunha, 1914 - 15

N	d 10°	20°	30°	40°	50°
0	0.99973	0.99823	0.99567	0.99224	0.98897
0.1	1.0143	1.0127	1.0100	1.0067	1.0028
0.2	.291	.274	.0245	.218	.0198
0.5	.730	.705	.0677	.641	.0604
0.8	.243	.213	.1187	.1153	.1120
1	.446	.413	.1384	.1352	.1315
1.5	2.095	2.063	.2033	.1995	.1946
2	2.767	2.734	.2689	.1648	.2593

N	η (water^t = 1) 10°	20°	30°	40°	50°
0.1	1.0036	0.9986	1.0050	1.0116	1.0177
0.2	1.0097	.9976	.0086	.0222	1.0312
0.5	1.0257	.9229	.0404	.0513	1.0653
0.8	1.0413	.9554	.0757	.0834	1.0928
1	1.0419	.9714	.0863	1.1024	1.0218
1.5	1.0808	1.1133	2.1389	1.1607	1.1883
2	1.1055	1.1586	2.1857	1.2176	1.2560

Gibson, 1931

mol%	π	d
	25°	
0	39.46	0.9970
10	33.39	1.0856
20	27.93	.1898
30	22.79	.3168
40	18.09	.4685
50	13.82	.6555

Water + Cesium selenate (Cs_2SeO_4)

Tutton, 1897

71.01% f.t. = 12°

%	d
20°	
45.94	1.5841
53.43	1.7432

H_2O + Cesium pentaborate (CsB_5O_8)

Rollet and Andres, 1931

f.t.	%	f.t.	%
-0.36	1.59 E	89.4	17.85
+5	1.68 (8+1)	93	19.5
18	2.33	93.6	19.75
30	3.52	95	20.4
45	5.57	99.1	22.3
60	8.31	99.5	22.6
75	12.0	101.65	23.45
80	13.8		

sat.sol. b.t. 102.00°

H_2O + Cesium formate ($CsHCO_2$)

Sidgwick and Gentle,1922

%	f.t.	%	f.t.
4.00	-0.77	88.67	42.7
8.11	-1.72	90.42	44.6
14.29	-3.44	-	45.0
19.62	-5.27	92.54	43.8
26.75	-8.12	93.47	42.6
77.07	+1.0	-	41.0
81.69	21.0	95.27	95.4
83.25	26.2	96.67	161.6
84.81	32.2	100	265
86.88	39.2		

(1+1)

H_2O + Cesium acetate ($CsC_2H_3O_2$)

Sidgwick and Gentle,1922

%	f.t.	%	f.t.
6.63	- 1.36	91.06	+ 21.5
19.38	- 9.32	91.98	61.1
25.26	- 7.92	93.09	88.6
34.22	-19.07	95.78	133.8
89.71	- 2.5	100	194

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.945	1.0	1.026
0.2	0.947	1.2	1.049
0.3	0.954	1.4	1.072
0.4	0.964	1.6	1.095
0.5	0.975	1.8	1.119
0.6	0.986	2.0	1.142
0.7	0.996	2.5	1.196
0.8	1.006	3.0	1.251
0.9	1.016	3.5	1.306

H_2O + Cesium benzoate ($CsC_7H_5O_2$)

Sidgwick and Ewbank, 1921

%	f.t.	%	f.t.
9.92	- 1.22	74.78	+ 12.0
25.62	- 4.44	77.34	53.5
43.08	-10.81	81.51	124.0

H_2O + Cesium salicylate ($CsC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
16.08	- 1.83	80.07	41.5(1+2)
30.97	- 4.24	85.67	54.0
47.71	- 8.08	89.91	67.5
65.70	- 3.05	(1+1) 92.80	84.3
68.24	+ 4.5	94.33	107.5
74.59	20.5		

H_2O + Cesium-m-oxybenzoate ($CsC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
16.98	- 1.97	81.70	29.0
32.84	- 4.92	84.45	35.0
48.53	- 9.65	86.35	40.0
63.15	-18.50	87.88	41.5
77.10	+10.0	89.88	75.0
78.52	19.0	93.92	126.5

(X aq. +1)

H_2O + Cesium-p-oxybenzoate ($CsC_7H_5O_3$)

Sidgwick and Ewbank, 1922

%	f.t.	%	f.t.
14.34	- 1.56	60.36	75.5
20.38	- 2.42	65.35	90.3
29.57	- 4.27	70.67	107.5
29.57	+ 3.5	73.84	118.0
40.80	27.5	79.88	136.0
50.07	49.0		

(1+1)

H_2O + Cesium telluroxalate ($Cs_2C_2H_6O_{20}Te$)

Rosenheim and Weinheber, 1911

%	f.t.	%	f.t.
6.03	0	16.52	40
11.02	20	21.67	50
13.10	30		

H_2O + Thallium fluoride (TlF)

Lübben, 1914

N	d	n 4600.9A°	4466.2	4229.5
		18°		
0	0.99862	1.33878	1.33992	1.34132
0.2006	1.04120	1.34256	1.34373	1.34518
0.5010	1.10447	1.34811	1.34932	1.35084
1.012	1.21129	1.35741	1.35865	1.36032
2.015	1.41774	1.37515	1.37669	1.37852

N	d	n 3848.9	3749.6	3611.9
0	0.99862	1.34477	1.34586	1.34755
0.2006	1.04120	1.34876	1.34990	1.35166
0.5010	1.10447	1.35463	1.35581	1.35767
1.012	1.21129	1.36440	1.36569	1.36774
2.015	1.41774	1.38326	1.38478	1.38711

N	d	n 3467.0	3403.6	3333.7
0	0.99862	1.34961	1.35061	1.35179
0.2006	1.04120	1.35380	1.35787	1.35607
0.5010	1.10447	1.35994	1.36104	1.36232
1.012	1.21129	1.37021	1.37143	1.37284
2.015	1.41774	1.38996	1.39142	1.39305

N	d	n 3255.8	3133.3	3085.1
0	0.99862	1.35322	1.35572	1.35681
0.2006	1.04120	1.35756	1.36018	1.36131
0.5010	1.10447	1.36391	1.36668	1.36790
1.012	1.21129	1.37458	1.37766	1.37892
2.015	1.41774	1.39508	1.39873	1.40032

N	d	n 2981.1	2881.1	2748.7
0	0.99862	1.35944	1.36224	1.36653
0.2006	1.04120	1.36406	1.36706	1.37154
0.5010	1.10447	1.37084	1.37401	1.37893
1.012	1.21129	1.38219	1.38572	1.39129
2.015	1.41774	1.40417	1.40841	1.41525

Heydweiller, 1921

N	d	λ
	18°	
0.1	1.02132	92.6
0.2	04246	86.8
0.5	1057	78.7
1	2102	71.5
2	4162	62.7

Jauch, 1921

N	U	N	U
	18°		
0.2	0.9485	1.5	0.7312
0.5	.8881	2	0.6712
1	.8021		

H_2O + Thallium chloride (TlCl)

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
4.2	144	26.4	338
6.67	177	30 - 86	364
9.0	205	81.5	367
10.0	213	92.2	368
12.3	234	96.2	381
21.1	303	100	430

Berkeley and Applebey, 1911

sat.sol. b.t. = 100.098

H_2O + Thallium bromide (TlBr)

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
1.73	162	7.25	294
3.09	215	89.5	414
4.45	258	92.0	421
6.2	209	100	457

H_2O + Thallium hydroxide (TlOH)

Bahr, 1911

N	d	f.t.
	15°	
1.151	1.231	0
1.554	1.317	18.5
1.582	1.322	19.5
1.803	1.342	29
1.861	1.377	32.1
1.967	1.400	33.1
2.075	1.417	36
2.240	1.446	40
2.442	-	44.5
2.940	-	54.1
3.281	-	59.4
3.601	-	64.6
4.673	-	78.5
5.705	-	90.0
6.708	-	99.2

Faust, 1927

N	d	η	σ
		20°	
1.4117	1.2923	680	74.54
0.696	1.1461	797	73.1
0.343	1.0734	887	72.8

H_2O + Thallium nitrate ($TlNO_3$)

Applebey and Hughes, 1915

t	p	t	p
sat.sol.		sat.sol.	
93.55	547.1	104.17	753.8
95.50	577.3	104.30	755.0
96.78	633.5	105.12	773.0
97.91	653.7	106.00	795.1
98.99	659.8	106.11	795.8
100.98	679.5	106.83	813.4
101.86	699.0	107.54	827.6
103.15	725.2	107.68	826.6
103.73	743.0	108.19	836.1
103.79	740.5	108.29	842.3
104.05	743.0	108.84	859.1

Etard, 1894

%	f.t.	%	f.t.
4.2	3.5	85.0	107
8.8	18	95.0	135
13.2	32	95.2	145
30.4	58	96.5	150
74.5	95	97.0	155

Berkeley, 1904

%	f.t.	%	f.t.
3.91	0.65	36.03	64.95
7.35	15.40	47.90	76.00
12.76	30.60	63.50	87.80
19.99	44.65	85.59	104.5
29.23	57.30		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.881	0.3	0.800
0.2	0.833	0.4	0.775

Berkeley, 1904

%	t	d
	sat.sol.	
3.91	0.65	1.0346
7.35	15.40	1.0653
12.76	30.60	1.1150
19.99	44.65	1.1891
29.23	57.30	1.2986
36.03	64.95	1.3957
47.90	76.00	1.6096
63.50	87.80	2.0258
85.59	104.5	3.1906

Rabinowitsch, 1921

%	d	η (water^t=1)	$\varkappa$
	100°		
7.32	2.292	1.918	5250
68.3	2.111	.702	5060
63.5	1.948	.560	4830
56.7	.756	.421	4460
49.2	.597	.308	3960
39.6	.423	.188	3340
29.8	.275	.115	2640
20.7	.155	.085	1910
14.1	.056	.039	1350
10.5	.50	.023	1040
5.43	.0016	1.005	570
2.08	0.9833	-	310
0	0.9	1.000	0

H_2O + Thallium carbonate (Tl_2CO_3)

Crookes, 1864

%	f.t.
15.55	3.87
60.00	10.47

H_2O + Thallium perchlorate ($TlClO_4$)

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.900	0.4	0.821
0.2	0.867	0.5	0.804
0.3	0.842		

Carlson, 1910

%	f.t.	d
5.66	0	1.060
7.44	10	1.075
16.47	30	1.146
28.38	50	1.251
39.44	70	1.430
44.90	80	1.520

H_2O + Thallium sulfate (Tl_2SO_4)

Morey and Chen, 1956

t	P kg
V + L + C	
374	above 280
400	235
500	277

Berkeley and Applebey, 1911

Sat.sol. b.t. = 100.317

Berkeley, 1904 and
Cohen, Ishikawa and Moesveld, 1923

%	f.t.	%	f.t.
2.64	0.15	9.87	60.40
4.14	15.60	12.15	75.90
5.77	29.80	14.11	90.05
7.74	30.00	15.57	99.70
	44.95		

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
17.7	116	225	31.2
18.8	130	233	32.2
19.0	131	241	32.4
21.5	146	249	32.5
22.4	153	271	34.4
28.8	167	272	34.8
25.5	173	359	37 - 76
27.7	186	365	78.8
27.0	188	376	84.0
29.0	208	380	85.7
29.6	209	632	100

H_2O + Thallium acetate ($TlC_2H_3O_2$)

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.913	1.6	0.814
0.2	0.891	1.8	0.810
0.3	0.876	2.0	0.807
0.4	0.865	2.5	0.801
0.5	0.855	3.0	0.796
0.6	0.849	3.5	0.789
0.7	0.843	4.0	0.783
0.8	0.838	4.5	0.777
0.9	0.833	5.0	0.772
1.0	0.829	5.5	0.766
1.2	0.823	6.0	0.760
1.4	0.818		

Guillaume, 1946

%	d	*(α)magn.10^6 5780Å	n
	18°		
27.2	1.2750	3.836	1.3658

*in radians, gauss, centim.

H_2O + Thallium ammonium tartrate ($TlC_4H_8O_6N$)

Long, 1880

%	d	$(\alpha)_D$
	20°	
0	0.9982	-
4.821	1.0354	10.032
9.306	1.0726	8.815
13.499	1.1092	7.914
17.42	1.1459	7.563

H_2O + Thallium lithium tartrate ($TlLiC_4H_4O_6$)

Long, 1809

%	d	$(\alpha)_D$
	20°	
0	0.9982	-
4.59	1.0347	9.456
6.05	1.0744	7.799
12.84	1.1107	7.139
16.57	1.1467	6.693

H_2O + Thallium sodium tartrate ($NaTlC_4H_4O_6$)

Long, 1880

	%	d	$(\alpha)_D$
		20°	
	4.06	1.0317	9.065
7.86	7.86	1.0652	7.701
	11.42	1.0992	6.992
	14.80	1.1311	6.492

H_2O + Thallium potassium tartrate ($KTlC_4H_4O_6$)

Long, 1889

%	d	$(\alpha)_D$
	20°	
4.81	1.0377	10.057
9.27	1.0764	8.840
13.44	1.1143	8.365
17.32	1.1530	8.173

H_2O + Silver fluoride (AgF)

Guntz and Guntz Jr., 1924

%	f.t.	%	f.t.
0	0	62.96	19 (2+1)
28.57	- 8	63.24	20
37.50	-14.2 E	64.03	24
46.58	0	64.24	25
54.55	+10	65.52	30
62.26	18.5	65.86	32
62.75	18.6	68.95	39.5 (2+1)-0 aq.
62.89	18.65(4+1)-(2+1)		
63.50	0 (5+3)	68.15	27 (5+3)-(1+1)
65.13	10	68.28	28.5 (1+1)
66.95	21.5	68.75	30
67.63	25	69.32	38.2 (1+1)-0 aq.

%	f.t.
68.95	39.5 (2+1)-0 aq.
68.55	50 0 aq.
67.21	108

Heydweiller, 1912

N	d	κ	N	d	κ
			18°		
0.1001	1.01140	80.8	0.8044	1.09866	478
0.2010	.02424	149.5	1.584	.19439	790
0.4034	.04922	271.8	3.136	.38016	1193

Guntz and Guntz Jr., 1914

%	d	%	d
	18°		
7.20	1.07	56.4	2.09
29.6	1.38	66.2	2.62
49.2	1.82		

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34583
0.1	1.33465	2.0	1.35647
0.2	1.33602	3.0	1.36703
0.5	1.33982		

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.8044	1.34360
0.1001	1.33464	1.584	1.35185
0.2010	1.33603	3.136	1.36838
0.4034	1.33859		

H_2O + Silver nitrate ($AgNO_3$)

Heterogeneous equilibria

Smits, 1904

t	p	t	p
	sat.sol.		
133	760	170	1010
135	800	185	900
150	960	191	760
160	1000		

Boswell and Cantelo, 1922

N	Dp/p_o	N	Dp/p_o
	23°		
6.000	0.114	2.000	.036
4.000	.072	1.000	.015

Dingemans and Van den Berg, 1942

t	p	t	p
	sat.sol.		
10.00	8.1	107.0	455
12.00	9.2	110.0	489
15.00	11.0	112.0	512
17.00	12.4	115.0	547
20.00	14.7	117.0	571
22.00	16.5	120.0	607
25.00	19.5	122.0	631
27.00	21.8	125.0	667
30.00	25.5	127.0	691
32.00	28.3	130.0	727
35.00	33.0	132.0	751
37.00	36.4	135.0	786
40.00	42.2	137.0	809
42.00	46.4	140.0	840
45.00	53.3	142.0	859
47.00	58.4	145.0	887
50.00	66.7	147.0	905
52.00	72.8	150.0	928
55.00	82.7	152.0	942
57.00	89.9	155.0	959
60.00	101.5	157.0	967
62.00	109.9	160.0	976
65.00	123.3	162.0	985
67.00	132.9	165.0	995
70.0	148.4	167.0	999
72.0	159.5	170.0	999
75.0	177.2	172.0	996
77.0	189.8	175.0	985
80.0	209.9	177.0	972
82.0	224.0	180.0	944
85.0	246.4	182.0	921
87.0	262.2	185.0	879
90.0	287.2	187.0	847
92.0	304.7	190.0	786
95.0	332.2	192.0	735
97.0	351.4	195.0	640
100.0	381	197.0	580
102.0	401	200.0	465
105.0	433	202.0	395

Campbell, Fishman and al., 1956

t	p	t	p
	9.98 %		
30.06	31.64	70.00	230.4
40.00	54.63	80.00	349.8
50.00	91.37	90.00	517.5
60.00	147.4	101.00	774.8
	19.98 %		
30.06	31.57	70.00	227.5
40.00	54.36	80.00	344.9
50.00	90.81	90.00	509.0
60.00	146.0	101.00	760.7
	30.00 %		
30.00	30.30	70.00	220.6
40.00	52.63	80.00	334.8
50.00	87.77	90.00	495.0
60.00	141.3	102.00	767.5
	40.00 %		
30.00	29.50	70.00	214.5
40.00	51.18	80.00	325.7
50.00	85.40	90.00	481.4
60.00	137.5	103.00	772.8
	50.00 %		
30.00	28.68	70.00	208.6
40.00	49.78	80.00	316.3
50.00	82.95	90.00	466.9
60.00	133.6	103.00	749.1
	59.29 %		
31.04	29.33	80.50	308.6
40.09	47.87	90.56	457.5
50.14	79.70	100.65	662.0
60.28	129.2	105.60	787.1
70.37	202.7		
	69.99 %		
30.00	26.15	70.00	189.6
40.00	45.88	80.00	286.2
50.00	75.79	90.00	421.5
60.00	122.0	100.00	606.6
	80.05 %		
60.00	107.2	85.00	307.9
65.00	134.2	90.00	372.4
70.00	166.8	95.00	447.9
75.00	206.2	100.00	534.8
80.00	252.6		
	85.07 %		
75.00	183.1	87.50	298.8
80.00	223.8	90.00	328.6
82.00	246.6	92.50	360.8
85.00	271.8	100.00	471.6
	90.00 %		
120.00	716.1	126.58	866.3
124.80	820.2	130.00	950.4

%	Q vap	%	Q vap
	50°		
9.98	10227	59.29	10192
19.98=	10172	69.99	10138
30.00	10204	80.05	10135
40.00	10202	85.07	10039
50.00	10192		

Robinson and Fait, 1941

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
	25°		
0.1250	0.1297	2.076	3.433
.1771	.1856	.108	.518
.1884	.1984	.371	4.253
.2765	.2958	.464	.492
.3303	.3563	.494	.602
.5056	.5635	.528	.758
.6580	.7620	.677	5.237
.7755	.9232	.724	.385
1.011	1.275	.984	6.278
.210	.605	3.014	.393
.462	2.057	.040	.472
.478	.072	.160	.917
.492	.097	.230	7.180
.572	.270	.443	8.000
.582	.287	.693	8.986
.943	3.066	.805	9.467
.958	.148	.985	10.18
.983	.208	4.271	11.25
2.013	.262	.540	12.41
.026	.274	.787	13.38
.043	.359	.811	13.48
.067	.394		

m_1 = m of KCl m_2 = m of $AgNO_3$

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.903	1.6	0.680
0.2	0.870	1.8	0.662
0.3	0.847	2.0	0.646
0.4	0.827	2.5	0.609
0.5	0.811	3.0	0.576
0.6	0.795	3.5	0.550
0.7	0.779	4.0	0.523
0.8	0.766	4.5	0.502
0.9	0.754	5.0	0.483
1.0	0.742	5.5	0.467
1.2	0.720	6.0	0.453
1.4	0.699		

Robinson and Stokes, 1949

m	activity coefficient	m	activity coefficient
	25°		
0.1	0.734	1.6	0.352
0.2	0.657	1.8	0.333
0.3	0.606	2.0	0.316
0.4	0.567	2.5	0.280
0.5	0.536	3.0	0.252
0.6	0.509	3.5	0.229
0.7	0.485	4.0	0.210
0.8	0.464	4.5	0.194
0.9	0.446	5.0	0.181
1.0	0.429	5.5	0.169
1.2	0.399	6.0	0.159
1.4	0.374		

Baroni, 1893

%	b.t.	%	b.t.
2.12	100.109	7.46	100.398
3.33	100.173	9.31	100.505
4.47	100.234	11.63	100.642
5.79	100.306		

Buchanan, 1899 and 1912

%	b.t.	%	b.t.
45.70	102.93	17.35	100.81
40.09	102.43	15.51	100.71
33.86	101.92	13.89	100.61
30.91	101.72	11.95	100.51
28.24	101.52	10.05	100.41
25.60	101.32	7.97	100.31
22.25	101.11	0.00	99.87
18.97	100.91		

Kahlenberg, 1901

%	b.t.	%	b.t.
3.78	100.20	7.68	100.43
10.65	100.58	14.01	100.78
19.02	101.09	20.09	101.10
26.36	101.56	23.44	101.56
31.46	101.96	31.22	101.93
36.20	102.30	35.20	102.22
40.72	102.70	39.47	102.57
45.46	103.12	42.75	102.86
49.94	103.60	46.36	103.22
52.43	103.84	49.93	103.57
55.02	104.15	53.54	103.98
57.67	104.50	57.71	104.58

Campbell and Kartzmark, 1950

%	b.t.	%	b.t.
	760 mm		
1.61	100.055	65.92	105.18
13.59	100.735	66.77	105.48
25.77	101.400	70.04	106.30
33.50	101.795	77.39	108.6
41.44	102.310	82.36	111.1
47.33	103.105	85.14	113.6
54.52	103.85	90.55	122.3
56.64	103.91	92.14	134.1

Kremers, 1854

%	f.t.	%	f.t.
54.95	0	87.72	85
69.44	19.5	91.01	110
83.33	54		

Rüdorff , 1872

%	f.t.	%	f.t.
3.85	-0.7	24.24	-4.10
7.41	-1.4	26.47	-4.55
9.09	-1.6	28.57	-4.85
10.71	-1.9	30.55	-5.1
13.79	-2.5	32.43	-5.3
16.67	-2.95	34.21	-5.6
21.88	-3.75		

Guthrie, 1876

%	f.t.	%	f.t.
10	- 0.8	50	- 5.5
20	- 2.7	53	- 2.2
30	- 4.7	55	0.0
40	- 6.0	69.4	+19.5
48.3	- 6.5 E		

Tilden and Shenstone, 1884

%	f.t.
94.19	125
95.10	133

Etard, 1894

%	f.t.	%	f.t.
46.2	- 7	75.7	36.5
46.0	- 7	76.9	40.5
47.6	- 5	77.1	45
52.4	- 1	78.5	48
51.9	- 1	84.0	73
56.3	+ 5	88.7	122
61.2	10	92.1	134
66.1	15.5	92.8	135
67.8	20	92.7	135
71.1	26	93.3	148
73.0	29	95.2	160
73.8	31	96.9	182

Greenish and Smith

65.4% f.t. = 15.56

Middelberg, 1903

E: 47.1% - 7.3°

Kasanzew, 1923

%	f.t.	%	f.t.
55.6	0	74.1	30
63.3	10	80.2	50
66.65	15	90.0	100
71.8	25		

Cohen, De Meester and Moesveld, 1924

73.8% f.t. = 30°

Campbell and Boyd, 1943

%	f.t.	%	f.t.
0	0	81.40	59.99
10.41	- 1.89	82.98	64.92
19.72	- 3.55	83.92	70.61
30.04	- 5.00	84.03	76.02
39.16	- 6.36	86.17	84.42
44.72	- 7.22	87.59	95.43
48.13	- 6.51	88.44	103.57
54.04	- 0.88	89.74	114.25
58.64	4.32	90.42	124.05
63.11	10.41	92.23	134.65
65.90	15.37	93.05	142.0
68.20	19.66	93.52	146.10
70.96	25.18	94.49	151.1
72.81	30.57	95.45	155.0
74.31	35.48	96.45	160.9
75.80	40.07	97.02	168.0
77.47	44.98	98.07	181.1
78.95	50.16	99.06	194.9
80.24	55.29		

tr.t.: 159°

E : 46.9% E: - 7.57°

Properties of phases .

Wiedemann, 1856

c	%	d
	16.6°	
0	0	0.9989
4.25	4.11	1.0335
5.33	5.09	1.0456
8.5	7.95	1.0679
12.75	11.57	1.1008
17.0	15.02	1.1308

Forster, 1878

%	t	d
0	24.0	0.9973
57.87	24.0	1.8652
29.01	23.0	1.3106
15.82	24.0	1.1497

Kohlrausch, 1879

%	d	%	d
	18°		
5	1.0422	35	1.3945
10	1.0893	40	1.4773
15	1.1404	45	1.5705
20	1.1958	50	1.6745
25	1.2555	55	1.7895
30	1.3213	60	1.9158

Kanonnikoff, 1885

%	t	d
21.3	18.6	1.2097
13.51	22.2	1.1224
0	20.0	0.9982

Kohlrausch, 1892

%	d	%	d
	15°		
0	0.9991	20	1.1969
5	1.044	40	1.4791
10	1.0901		

Sentis, 1897

mol%	t	d
10	21.1	1.6858
8	19.7	1.5748
5	20.3	1.3641
3	19.2	1.2236
2	18.0	1.1508
1	17.7	1.0759
1	19.8	1.0754

Chéneveau, 1907

%	d	%	d
	19°		
20.84	1.2030	9.25	1.0816
17.25	1.1632	4.82	1.0408
13.39	1.1228	0	0.9984
11.37	1.1019		

Bousfield and Lowry, 1910

t	d			
	70%	10%	2%	0%
7	-	1.09129	1.01731	0.99993
15	2.23812	1.08963	1.01635	0.99913
25	2.22831	1.08662	1.01913	0.99707
40	2.21312	1.08070	1.00913	0.99224
60	2.19235	1.07055	0.99993	0.98324
80	2.17091	1.05845	0.98850	0.97384

Sakhanov, 1913, Sakhanov and Rabinovich, 1915

%	d	%	d
	25°		
0	0.9971	49.81	1.6548
20.04	1.1920	67.38	2.1335
34.32	1.3803		

Rabinovich, 1921

%	d	%	d
	99.9°		
86.9	2.956	34.1	1.324
79.9	2.567	26.5	1.213
72.5	2.246	18.5	1.127
63.1	1.937	9.32	1.037
53.3	1.738	3.20	0.9835
49.0	1.571	1.57	0.9698
41.5	1.436		

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.0703	3	1.4131
1	1.1395	4	1.5479
2	1.2768		

Goard, 1925

N	d	N	d
	20°		
1	1.118	3	1.348
2	1.236	5	1.402

Rehbinder, 1926

%	d	%	d
	100°		
90.39	3.195	44.9	1.445
81.7	2.657	33.2	1.310
81.5	2.646	22.6	1.170
79.1	2.525	16.2	1.098
69.1	2.125	0	0.9584
51.2	1.622		

Skanavi, Grigoryeva and Shternin, 1935

%	d	%	d
	25°		
76.56	1.9968	60.18	1.6650
75.52	1.9644	49.30	1.4900
66.93	1.7992	41.86	1.3910

76.56%			
d			
30°	35°	40°	50°
1.9930	1.9899	1.9861	1.9757

Jones and Colvin, 1940

N	d	N	d
	25°		
0.002	0.99736	0.20	1.02518
.005	.99778	.50	.06709
.01	.99883	.99719	.13601
.015	.99919	2.0	.27363
.02	.99992	3.00425	.41002
.05	1.00414	4.0115	.54565
.10	.01118	4.99808	.67745
	0°		
0.001	0.99991	0.2	1.02880
.002	1.00016	0.5	1.07174
.005	.00059	1.00326	1.14293
.01	.00132	2.0157	1.28381
.015	.00204	3.03205	1.42306
.02	.00278	4.05174	1.56115
.05	.00714	5.05	1.69487
.1	.01438		

Guillaume, 1946

%	t	d
17.85	21	1.1719
67.40	16	2.1549

Campbell and Kartzmark, 1950

%	N	d
	24.99°	
1.68	0.100	1.011
15.01	1.004	1.137
26.74	1.998	1.276
36.41	3.028	1.413
44.02	4.000	1.544
50.71	5.029	1.685
59.28	6.006	1.810
61.30	7.012	1.943
65.61	8.011	2.075
69.45	9.010	2.203
71.59	9.709	2.298

%	mol %	d
	25°	
1.68	0.178	1.0111
14.96	1.83	1.136
26.68	3.71	1.273
36.22	5.67	1.407
44.08	7.71	1.542
50.58	9.78	1.675
56.16	11.95	1.806
61.07	14.25	1.938
65.49	16.74	2.069
71.59	21.07	2.298

Campbell and Kartzmark, 1952

%	d	%	d
	95°		
0.9363	0.9696	60.46	1.852
18.572	1.116	69.54	2.157
29.70	1.252	73.38	2.293
36.954	1.364	79.45	2.540
50.735	1.617	82.33	2.816

Campbell, Bisset and Bednas, 1953

%	d	%	d
	35°		
0.0	0.99406	45.764	1.5705
1.6612	1.0082	51.422	.6874
20.051	1882	51.844	.6970
24.350	2389	57.194	.8250
36.551	4094	64.283	2.0253
42.849	5158		

N	d	N	d
	25°		
0.0000	0.99707	3.8399	5225
0.0989	1.0112	5.1313	1.6953
1.4077	1928	5.2027	7049
1.7830	2440	6.1730	8336
3.0458	4157	7.6996	2.0349

Campbell, Kartzmark and al., 1954

%	N	d
221.7°		
1.954	0.0965	0.8393
18.94	0.119	1.003
28.56	1.868	1.111
41.03	3.130	1.296
48.65	4.047	1.413
48.84	4.076	1.418
54.94	4.900	1.515
59.27	5.763	1.652
60.91	6.159	1.718
65.84	7.097	1.831
69.63	8.124	1.982
75.40	9.761	2.199
81.56	11.98	2.495
85.92	14.32	2.832
88.91	16.01	3.059
92.06	17.71	3.268
100	23.19	3.94

Wiedemann, 1856

c	%	η
	16.6°	
0	0	1095
4.25	4.11	1134
5.33	5.09	1159
8.5	7.95	1190
12.75	11.57	1238
17.0	15.02	1276

Sakhanov, 1913

%	η	%	η
	25°		
0	895	49.81	1451
20.04	990	67.38	2439
34.32	1134		

Sakhanov and Rabinovich, 1915

%	η	%	η
	25°		
0	891	49.81	1451
20.04	990	67.38	2439
34.32	1134		

Rabinovich, 1921

%	η (water=1)	%	η (water=1)
	99.9°		
86.9	8.59	34.1	1.40
79.9	5.12	26.5	1.25
72.5	3.57	18.5	1.134
63.1	2.63	9.32	1.033
53.3	2.17	3.20	1.012
49.0	1.84	1.57	1.010
41.5	1.59	0	1.000

Jones and Colvin, 1940

N	η(water=1)	N	η(water=1)
	25°		
0.002	1.00036	0.20	1.01224
.005	.00068	0.50	.03088
.01	.00110	0.99719	.06783
.015	.00148	2.0	.16788
.02	.00177	3.00425	.29951
.05	.00364	4.0115	.46471
.10	.00652	4.99808	.66223
	0°		
0.001	1.00013	0.2	0.99701
.002	.00017	0.5	0.99782
.005	.00021	1.00326	1.01393
.01	.00018	2.0157	1.08576
.015	.00007	3.03205	1.19842
.02	1.00000	4.05174	1.34860
.05	0.99939	5.05	1.53375
.1	0.99847		

Campbell and Kartzmark, 1950

%	N	η(water=1)
	24.99°	
1.68	0.100	1.009
15.01	1.004	1.074
26.74	1.998	1.177
36.41	3.028	1.307
44.02	4.000	1.472
50.71	5.029	1.658
59.28	6.006	1.890
61.30	7.012	2.170
65.61	8.011	2.505
69.45	9.010	2.998
71.59	9.709	3.330

Campbell and Kartzmark, 1952

%	η(water=1)	%	η(water=1)
	95°		
0.9363	0.9928	60.46	2.280
18.572	1.122	69.54	3.107
29.70	1.263	73.38	3.590
36.954	1.425	79.45	4.262
50.735	1.797		

Campbell, Bisset and Bednas, 1953

%	η	%	η
	35°		
0.0	722.5	45.764	1107.4
1.6612	728.9	51.422	1240.1
20.051	810.0	51.844	1253.3
24.350	838.5	57.194	1428.1
36.551	958.4	64.283	1778.6
42.849	1055.7		
	25°		
0.0000	893.7	4.2526	1342.7
0.0989	900.5	5.1313	1503.1
1.4077	987.1	5.2027	1519.8
1.7830	1020.8	6.1730	1734.2
3.0458	1163.0	7.6996	2169.3
3.8399	1277.0		

Sentis, 1897

mol%	t	σ
10	21.1	78.7
8	19.7	78.0
5	20.3	76.5
3	19.2	75.6
2	18.0	74.9
1	17.7	74.4
1	19.8	73.7
0	13.5	74.0
0	25.1	72.3

Zemplen, 1906

t	σ	t	σ
0.074%		0.148%	
35.2	72.5606	34.6	72.9521
61.4	67.2907	61.3	66.9856
94.3	62.9390	95.3	63.4946
1.484%		14.84%	
34.6	69.9355	34.6	78.2289
60.7	67.1053	61.7	73.8497
94.0	62.8350	91.9	68.6030
27.92%		61.56%	
34.6	85.1439	34.3	90.5208
60.4	80.1968	61.4	86.1828
91.3	72.8137	91.0	82.8904

Goard, 1925

N	σ	N	σ
	20°		
1	74.37	3	75.94
2	75.23	5	77.69

Rehbinder, 1926

%	σ	%	σ
	100°		
9.61	101.0	55.1	64.4
18.3	79.7	66.8	62.4
18.5	79.6	77.4	61.0
20.9	77.6	83.8	60.1
30.9	72.7	100.0	58.7
48.8	65.9		

Forster, 1878

%	t	n_D
57.87	24.0	1.43650
29.01	24.0	1.37044
15.82	23.0	1.35093
0	24.0	1.33256

Kanonnikoff, 1885

%	t	n H_α	n D	n H_β
21.3	18.6	1.357406	1.359720	1.364440
13.51	22.2	1.346437	1.348406	1.353187
0	20.0	1.33130	1.33310	1.33738

Chéneveau, 1907

%	n_D	%	n_D
	19°		
20.84	1.3581	9.25	1.3433
17.25	1.3532	4.82	1.3382
13.39	1.3483	0	1.3331
11.37	1.3457		

Okazaki, 1933

%	Verdet's constant (3632A°)	%	Verdet's constant(3632 A°
	28°		
10.35	0.03989	31.43	.04332
21.74	.04160	40.71	.04515

Guillaume, 1946

%	t	n 5780 Å	* (α) magn.10^6
17.85	21	1.3560	3.578
67.40	16	1.4644	2.435

* in radians, gauss, centim.

Kohlrausch, 1879

%	κ	$\tau.10^4$
	18°	
5	254	219
10	473	218
15	678	216
20	866	213
25	1052	211
30	1231	210
35	1398	208
40	1554	206
45	1704	205
50	1841	205
55	1970	207
60	2086	210

Sloan, 1910 (fig.)

N	λ
	0°
4.70	25.70
3.76	28.76
1.88	37.22
0.95	44.46
0.48	50.56
0.24	55.40
0.12	59.75
0.06	63.10

Sakhanov, 1913 and Sakhanov and Rabinovich, 1915

%	λ	ϰ
	25°	
20.04	71.27	1002
34.32	56.88	1586
49.81	43.84	2127
67.38	29.52	2498

Rabinovich, 1921

%	ϰ	%	ϰ
	99.9°		
86.9	5530	34.1	3700
79.9	6130	26.5	3060
72.5	6240	18.5	2300
63.1	5890	9.32	1290
53.3	5490	3.20	500
49.0	4950	1.57	260
41.5	4390	0	0

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	77.5	3	48.26
1	67.4	4	42.27
2	55.18		

Campbell and Kartzmark, 1950

%	N	ϰ	%	N	ϰ
		24.99°			
1.68	0.100	109.1	59.28	6.006	2316
15.01	1.004	781.5	61.30	7.012	2413
26.74	1.998	1303	65.61	8.011	2501
36.41	3.028	1655	69.45	9.010	2522
44.02	4.000	1940	71.59	9.709	2523
50.71	5.029	2168			

Campbell and Kartzmark, 1952

%	ϰ	λ
	95°	
0.9363	159	298.3
18.572	2107	172.7
29.70	3236	147.8
36.954	3881	130.8
50.735	4979	103.1
60.46	5595	84.88
69.54	6015	68.12
73.38	6071	61.28
79.45	5982	50.35
82.33	5620	40.08

Campbell, Bisset and Bednas, 1953

%	ϰ	%	ϰ
	35°		
1.6612	1306	45.764	2372
20.051	1197	51.422	2590
24.350	1413	51.844	2606
36.551	1983	57.194	2782
42.849	2252	64.283	2954

Campbell, Kartzmark and al., 1954

%	N	ϰ
	221.7°	
1.954	0.0965	5615
18.94	1.119	3639
28.56	1.868	5286
41.03	3.130	7110
48.65	4.047	-
48.84	4.076	8030
54.94	4.900	8809
59.27	5.763	9420
60.91	6.159	9698
65.84	7.097	10230
69.63	8.124	10530
75.40	9.761	11050
81.56	11.98	11350
85.92	14.32	11220
88.91	16.01	10840
92.06	17.71	10190
100	23.19	7190

Heat constants

Kapustinskii, Yakuchevskii and Drakin, 1953

%	U	%	U
	at room temp.		
0	(0.9980)	16.03	0.8413
4.57	0.9509	19.36	0.8110
8.63	0.9100	23.45	0.7759
12.49	0.8741		

Campbell, Fishmann and al., 1956

%	Q vap	%	Q vap
	50°		
9.90	10227	59.29	10192
19.98	10172	69.99	10138
30.00	10204	80.05	10135
40.00	10202	85.07	10039
50.00	10192		

H_2O + Silver perchlorate ($AgClO_4$)

Hill, 1922

%	f.t.	%	f.t.
100	480	73.9	-58
88.8	99	70.4	-40
88.1	79	60.3	-24
87.2	50	45.2	-10
86.5	43 tr.t.	26.55	- 3
84.5	25 (1+1)	0.96	- 0.16
81.3	0	0	0

Smith and Ring, 1937

%	f.t.	%	f.t.
82.07	0	84.78	25
82.88	10	85.59	30
84.04	20	86.21	35

Hill, 1922

%	d	%	d
	15°		
88.8	3.069	84.5	2.806
88.1	3.022	81.3	2.667
87.2	2.995	73.9	2.315

Mazzucchelli and Rossi, 1927

%	d	
	15°	25°
4.99	1.04067	1.03816
9.98	1.08635	1.08317
14.29	1.12857	1.12491
15.03	1.13612	1.13238

Hantzsch and Düringen, 1928

%	n_D	d
	20°	
26.58	1.35873	1.2544
18.03	1.34932	1.1589
6.544	1.33839	1.04892
0	1.33300	0.99540

Smith and Ring, 1937

%	d	%	d
82.07	2.7251	84.78	2.8487
82.88	2.7594	85.59	2.8825
84.04	2.8163	86.21	2.9173

H_2O + Silver ammonium nitrate ($AgH_4O_6N_3$)

Skanavi, Grigoryeva and Shternin, 1935

%	d	%	d
	25°		
76.56	1.9968	60.18	1.6650
75.52	1.9644	49.30	1.4900
66.93	1.7992	41.86	1.3910

t	d	t	d
	76.56 %		
30	1.9930	40	1.9861
35	1.9899	50	1.9757

%	η(water=1)	%	η(water=1)
	25°		
76.56	3.648	60.18	1.787
75.52	3.135	49.30	1.462
66.93	2.178	41.86	1.311

t	η	t	η
	76.56 %		
30	3.3084	40	2.7726
35	3.0207	50	2.3708

%	ϰ			
	103°	109°	110°	112°
100	-	-	1340	1420
99.84	-	-	1380	1460
97.13	1500	1640	1670	1710
94.03	1520	1680	1710	1770
90.24	3100	3280	3340	3400
87.75	3960	4240	4310	4440
86.32	4340	4540	4580	4650
85.13	4610	4870	4910	5000
80.31	5790	6060	6600	6180
77.05	6480	6730	6760	6820
76.56	6690	7010	7090	7170
72.93	7000	-	-	-
72.77	-	-	7180	-
60.13	7060	7350	-	-

%	χ	%	χ
	25°		
76.56	0.276	60.18	0.322
75.52	0.289	49.30	0.295
66.93	0.317	41.86	0.266

t	χ	t	χ
	76.56 %		
30	0.301	69	0.481
35	0.327	70	0.485
40	0.352	71	0.489
50	0.400	72	0.493
68	0.476		

H_2O + Silver Thallium nitrate ($AgTlN_2O_6$)

Rabinovich, 1938

%	d	κ	η (water=1)
	99.9°		
100	4.658	1150	75.3
98.6	4.440	1810	38.8
95.6	4.027	3070	16.6
91.0	3.526	4370	-
79.3	2.649	5610	3.47
79.4	2.296	5510	2.38
60.7	1.880	5070	1.815
50.7	1.639	4450	1.919
39.6	1.436	3770	1.326
30.7	1.283	2950	1.207
21.0	1.155	2150	1.109
10.5	1.049	1200	1.029
4.4	0.9924	950	1.010
1.4	0.9663	250	1.001

H_2O + Mercurous perchlorate ($HgClO_4$)

Newbery, 1936

%	f.t.	%	f.t.
68.2	-20	81.9	37
73.9	0	81.9	38
75.9	+ 8	82.0	40
78.0	17	82.7	50
78.9	21	83.3	60
80.7	30	83.7	70
81.8	35	84.2	78
82.2	36	85.2	99

P. WATER + SALTS OF POLYVALENT METALS .

L. WATER + SALTS OF Be, Mg, Sn and Pb .

Water + Beryllium chloride ($BeCl_2$)

Tammann, 1885

%	p	%	p
	100°		
2.86	748.0	16.83	606.7
5.65	731.0	18.90	573.7
8.98	704.6	23.18	492.5
12.35	668.8	25.82	440.6
15.29	629.4	28.34	295.9

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.01309	58.4
1	.02595	51.6
2	.0511	42.03
3	.0761	34.42
4	.1008	29

Jauch, 1921

N	U
18°	
0.5	0.9701
1	.9432
2	.8928
3	.8443
4	.7966

Water + Beryllium bromide ($BeBr_2$)

Tammann, 1885

%	p	%	p
	100°		
11.16	729.3	29.25	594.2
13.10	719.6	34.97	519.6
20.28	676.4	41.61	410.7
25.58	633.1	43.70	372.4

H_2O + Silver Thallium nitrate ($AgTlN_2O_6$)

Rabinovich, 1938

%	d	κ	η (water=1)
	99.9°		
100	4.658	1150	75.3
98.6	4.440	1810	38.8
95.6	4.027	3070	16.6
91.0	3.526	4370	-
79.3	2.649	5610	3.47
79.4	2.296	5510	2.38
60.7	1.880	5070	1.815
50.7	1.639	4450	1.919
39.6	1.436	3770	1.326
30.7	1.283	2950	1.207
21.0	1.155	2150	1.109
10.5	1.049	1200	1.029
4.4	0.9924	950	1.010
1.4	0.9663	250	1.001

H_2O + Mercurous perchlorate ($HgClO_4$)

Newbery, 1936

%	f.t.	%	f.t.
68.2	-20	81.9	37
73.9	0	81.9	38
75.9	+ 8	82.0	40
78.0	17	82.7	50
78.9	21	83.3	60
80.7	30	83.7	70
81.8	35	84.2	78
82.2	36	85.2	99

P. WATER + SALTS OF POLYVALENT METALS .

L. WATER + SALTS OF Be, Mg, Sn and Pb .

Water + Beryllium chloride ($BeCl_2$)

Tammann, 1885

%	p	%	p
	100°		
2.86	748.0	16.83	606.7
5.65	731.0	18.90	573.7
8.98	704.6	23.18	492.5
12.35	668.8	25.82	440.6
15.29	629.4	28.34	295.9

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.01309	58.4
1	.02595	51.6
2	.0511	42.03
3	.0761	34.42
4	.1008	29

Jauch, 1921

N	U
18°	
0.5	0.9701
1	.9432
2	.8928
3	.8443
4	.7966

Water + Beryllium bromide ($BeBr_2$)

Tammann, 1885

%	p	%	p
	100°		
11.16	729.3	29.25	594.2
13.10	719.6	34.97	519.6
20.28	676.4	41.61	410.7
25.58	633.1	43.70	372.4

Water + Beryllium iodate (BeI_2O_6)

Biber, Neiman and Bragina, 1941

%	f.t.	%	f.t.
11.08	0	19.13	30
12.34	10	25.62	40
14.31	15	23.80	50
15.55	20	42.33	60
17.50	25		

(4+1)

%	d	%	d
	15°		
1.30	1.0104	9.30	1.0875
1.92	.0155	10.35	.0995
2.90	.0251	13.37	.1287
4.03	.0342	14.31	.1422
5.23	.0483	15.63	.1555
6.61	.0603	17.36	.1756
8.21	.0713	21.38	.2262

Water + Beryllium nitrate (BeN_2O_6)

Sieverts and Petzold, 1933

%	f.t.	%	f.t.
13.0	-6.3	37.7	-59.0
12.6	-6.4	46.2(4+1)	-40.0
20.9	-15.3	49.4	+0.4
24.8	-23.0	51.2	15.0
27.5	-27.3	52.3	30.0
28.3	-29.7	58.6	50.0
30.9	-34.4	64.8	61.0
34.0	-46.0		

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.02024	63.8
1	.04009	56.2
2	.0800	45.80
3	.1197	37.5
4	.1592	30.84
5	.1984	24.85

Jauch, 1921

N	U
18°	
0.5	0.9619
1	.9287
2	.8664
3	.8149
4	.7698

Water + Beryllium perchlorate ($BeCl_2O_8$)

Lilich and Dzhurinskii, 1956

m	f.t.	m	f.t.
8.02	0	8.29	30
8.08	5	8.35	35
8.10	10	8.45	40
8.16	15	8.48	45
8.20	20	8.52	50
8.26	25		

Lilich and Mogilev, 1956

m	pH	m	pH
	25°		
		0.251	2.40
0.00217	3.95	0.594	2.04
.0183	3.35	1.00	1.81
.0726	2.91	3.76	0.97
.135	2.57		

H_2O + Beryllium periodade (BeI_2O_8)

Biber, Neiman and Bragina , 1941

%	f.t.	%	f.t.
44.86	0	55.33	15
48.97	5	56.41	20
52.01	10	57.84	25

(8+1)

Water + Beryllium sulfate ($BeSO_4$)

Tammann, 1885

t	p 0%	11.86%	21.80%	25.22%	30.79%
19.87	17.4	-	16.1	15.7	14.7
31.41	34.5	-	31.9	30.3	29.2
58.77	51.8	50.6	48.1	45.9	43.8
46.11	76.1	74.0	71.6	68.6	65.0
48.39	85.4	83.0	79.9	77.4	72.9
56.41	126.3	123.2	118.7	114.3	108.8
59.03	142.9	139.1	135.0	129.3	123.5
61.93	163.3	159.8	154.5	148.2	141.9
66.44	200.1	195.7	189.9	182.2	174.4
69.50	228.8	223.7	216.9	208.1	199.6
72.04	255.3	249.3	242.7	232.7	223.9
75.57	296.3	289.0	281.2	270.5	259.7
79.45	347.6	339.5	330.3	318.2	305.9
83.06	401.9	393.9	383.5	369.0	356.0
85.64	444.8	436.6	435.5	409.6	394.7
88.74	502.3	491.2	478.7	461.6	444.6
90.73	540.7	529.1	516.2	498.0	479.8
93.04	589.6	-	563.7	544.0	524.0
95.15	637.6	-	609.8	588.4	566.8
100.55	775.0	759.4	740.0	717.1	692.5

Tammann, 1885

%	p	%	p
	100°		
10.21	747.2	27.53	689.4
16.84	734.1	39.18	593.4
23.09	712.7	39.32	591.2

Robinson, 1952

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1022	0.06323	1.499	1.0510
0.1224	0.07510	1.674	1.228
0.1304	0.0804	1.888	1.508
0.2145	0.1287	2.177	1.791
0.3137	0.1844	2.194	1.808
0.3710	0.2157	2.205	1.824
0.4462	0.2615	2.457	2.146
0.5342	0.3091	2.497	2.206
0.5671	0.3298	2.589	2.328
0.5760	0.3362	2.678	2.456
0.6621	0.3922	3.039	2.995
0.6994	0.4162	3.064	3.041
0.7018	0.4174	3.135	3.147
0.8275	0.5011	3.229	3.298
0.8944	0.5485	3.429	3.615
0.9591	0.5988	3.506	3.750
1.128	0.7268	3.898	4.406
1.272	0.8458	4.086	4.730
1.280	0.8531	4.286	5.082

m_1 = m of $BeSO_4$ m_2 = m of NaCl

Levi-Malvano, 1906

mol%	f.t.	mol%	f.t.	mol%	f.t.
8.94	31 (6+1)	7.50	30 (4+1)	14.51	80 (2+1)
10.40	50	8.01	40	16.75	91.4
12.84	72.2	10.62	68	20.28	105
14.04	77.4	13.07	85	25.58	119
		15.53	95.4		
		19.76	107.2		
		21.98	111		

Parsons, Robinson and Fuller, 1907

%	f.t.	%	f.t.
0	0	4.85	-1.033
0.39	-0.103	6.02	-1.283
0.87	-0.206	8.69	-2.000
1.37	-0.305	9.93	-2.390
1.74	-0.374	10.83	-2.780
2.31	-0.474	12.77	-3.525
3.83	-0.772	14.23	-4.285

Schreiner and Sieverts, 1935

%	f.t.	%	f.t.
13.31	- 4.7	20.69	- 17.9
14.50	- 5.8	21.38	- 17.9
15.25	- 6.3	21.36	- 18.5
16.11	- 7.6	21.00	- 18.5
19.96	- 7.8	21.33	- 18.9
17.73	- 9.8	21.22	- 19.4 E
17.75	-10.2	21.00	- 18.0
18.16	-12.4	21.26	- 18.0
19.55	-13.3	21.15	+ 0.4 (4+1)
19.55	-14.3	22.30	+ 25.0
24.53	-15.2	23.79	+ 40.0
20.12	-15.6	25.87	+ 60.0
20.93	-16.4	28.72	+ 80.0
20.64	-16.6	31.17	+100.0
20.90	-16.8	31.97	+110.0

Jahn, 1891

c	d
20°	
0	0.9982
12.340	1.0936
23.906	1.1787

Gibson, 1934

mol%	d	mol%	d
	25°		
0	0.9970	15	1.1373
5	1.0416	20	1.1891
10	1.0883	25	1.2407

Gibson, 1934

mol %	π (1-1000 bars)	mol %	π (1-1000 bars)
0	39.46	15	26.58
5	34.81	20	23.08
10	30.62	25	20.04

Jahn, 1891

c	n H_α	n D	n H_β
		20°	
0	1.3315	1.3332	1.3375
12.340	1.3503	1.3522	1.3566
23.906	1.3660	1.3681	1.3726

c	(α) magn.
20°	
0	1
12.340	0.28526
23.906	0.29264

Marignac, 1876

%	U
21-52°	
2.84	0.9703
5.52	.9457
10.46	.9009
18.95	.8285

Kapustinskii, Yakuchevskii and Drakin, 1953

%	U	%	U
0	0.9980	13.57	0.8491
5.30	.9320	16.18	.8299
7.54	.9070	18.89	.8072
10.76	.8761		

Kapustinskii and Ruzavin, 1955

thermic conductivity coefficient $k.10^6$

%	k	%	k
		25°	
3.3	1434	12.5	1383
5.02	1427	19.3	1339
8.29	1407	23.3	1312
5.03	1431	15.5	1369
10.2	1399	23.8	1315

Water + Magnesium chloride ($MgCl_2$)

Heterogeneous equilibria

Tammann, 1885

t	p 0%	p 9.70%	p 15.98%	p 19.42%	p 25.71%
27.78	28.0	-	24.0	22.3	17.1
36.77	46.5	43.0	40.1	36.8	29.5
39.88	55.0	51.5	47.5	43.8	35.0
43.05	65.0	60.5	55.6	51.7	61.5
45.83	75.0	70.2	64.1	59.2	47.4
48.46	85.7	80.4	73.5	67.9	53.8
51.12	97.8	91.5	84.2	77.8	62.3
53.86	111.8	105.1	96.6	89.6	72.2
56.34	125.9	118.7	109.0	100.8	81.9
58.50	139.4	131.1	120.3	111.5	90.1
60.66	154.1	144.9	132.0	123.0	98.5
63.53	175.6	164.5	151.3	140.7	112.8
67.95	213.8	201.3	184.1	171.5	138.7
70.78	241.8	228.0	208.9	194.4	157.0
73.40	270.4	254.5	233.7	218.0	176.8
75.87	299.8	282.4	248.9	241.6	196.0
78.10	329.0	309.7	284.6	264.1	269.9
81.84	383.4	360.1	331.8	309.3	252.5
84.16	419.7	395.2	363.6	339.3	277.2
86.70	463.5	436.4	398.0	374.8	307.3
88.53	497.4	468.4	431.2	403.3	329.7
90.67	539.5	509.0	469.6	438.7	359.7
93.16	592.4	557.5	514.7	480.7	395.1
96.48	669.3	631.6	582.5	545.0	449.1
100.08	762.1	717.0	663.5	521.2	614.7

%	p	%	p
	100°		
6.02	736.1	26.51	501.9
9.29	717.4	29.75	435.3
14.76	675.2	31.52	396.7
19.25	619.4	34.07	339.7
23.12	561.6	37.21	282.7

Lescoeur, 1894

t	p dissoc. (4+1)	p dissoc. (6+1)	p sat.sol.
15	-	-	4.4
16	-	-	3.7
20	-	-	5.75
25	-	-	7.5
37.5	-	2	18
64.5	-	10	53
77.5	-	23	88
100	30	83	159
110	-	124	-
120	45	206	-
130	68	275	-

Van't Hoff and Meyerhoffer, 1898

p = 7.7 mm 25° sat. sol.

Speranski, 1909

t	p	t	p
	sat.sol.		
0	1.34	15	4.00
5	1.96	20	5.60
10	2.82	25	7.76

Derby and Yngve, 1915

t	p sat.sol.	p dissoc. (6+1) - (4+1)
	(6+1)	
10.14	3.4	-
18.05	5.2	-
25.53	7.7	-
31.60	-	1.6
31.65	11.2	-
41.92	18.5	-
42.10	-	2.7
44.46	21.4	-
50.96	29.0	-
55.19	35.4	-
58.31	-	7.8
64.10	-	11.3
64.99	54.1	-
68.27	-	13.2
69.38	-	14.6
74.27	-	21.3
76.53	-	21.9
88.85	124.2	41.5
96.96	154.1	63.1
99.23	162.7	-
99.43	163.6	-
104.7	189.9	98.5
108.2	-	116.8
109.9	195.3	-
110.2	198.6	-
110.9	197.6	-
112.9	198.3	-
113.1	197.4	146.1
113.6	196.0	-
114.9	193.7	-
115.6	-	165.4
116.0	191.6	-
116.6	184.0	-
117.2	169.1	-

t	p	t	p
	(4+1)		
117.7	170.8	122.6	205.0
118.0	173.7	125.5	225.3
119.2	181.5	128.9	247.9
121.2	193.1	138.3	307.1

Pohle, 1927

t	p	t	p
	sat.sol.		
30.0	10.5	60.0	43.0
35.0	14.0	65.0	58.0
40.0	18.0	70.0	69.0
45.0	24.0	75.0	82.0
50.0	29.0	80.0	102.0
55.0	35.0		

Kondiriev and Berezovski, 1935

t	p	t	p	t	p dissoc.
(6+1) sat.sol.		(6+1)-(4+1)		(4+1)	
15	4.11	20	0.67	10	0.10
20	5.80	30	1.40	20	0.30
30	9.95	40	2.51	30	0.55
40	17.33	50	4.36	40	1.15
50	28.62	60	8.06	80	3.04
60	44.33	80	24.93	100	11.40
70	65.52	100	74.86	116.7	28.64
80	92.21	116.7	156.92	117.2	29.20
90	127.40	150	432.0	150	148.30
100	165.06			160	228.62

Robinson and Stokes, 1940

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1120	0.1579	0.1587	0.2271
0.2656	0.3936	0.2795	0.4156
0.3644	0.5548	0.3924	0.6035
0.4815	0.7559	0.4840	0.7617
0.6503	1.075	0.7212	1.217
1.059	1.970	1.266	2.487
1.713	3.732	1.784	3.962
0.1917	0.2782	0.1944	0.2808
0.2933	0.4379	0.3492	0.5299
0.4040	0.6242	0.4387	0.6818
0.5293	0.8516	0.5600	0.9033
0.9856	1.798	0.9906	1.806
1.367	2.765	1.381	2.805
2.016	4.712		

m_1 = m of $MgCl_2$ m_2 = m of KCl

Stokes, 1945

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
1.143	2.175	1.743	3.833
1.645	3.541	2.050	4.810
1.661	3.591		

m_1 = m of $MgCl_2$ m_2 = m of KCl

Stokes, 1945

Isopiestic solutions at 25°			
m_1	m_2	m_1	m_2
2.168	3.327	3.843	6.307
2.480	3.860	4.026	6.658
2.782	4.379	4.373	7.345
2.829	4.463	4.502	7.608
3.111	4.959	4.893	8.427
3.136	5.004	4.982	8.615
3.287	5.277	5.476	9.740
3.371	5.426	5.840	10.61
3.809	6.245	5.925	10.84

m_1 = m of $MgCl_2$ m_2 = m of SO_4H_2

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.861	1.4	1.264
0.2	0.877	1.6	1.347
0.3	0.895	1.8	1.434
0.4	0.919	2.0	1.523
0.5	0.947	2.5	1.762
0.6	0.976	3.0	2.010
0.7	1.004	3.5	2.264
0.8	1.036	4.0	2.521
0.9	1.071	4.5	2.783
1.0	1.108	5.0	3.048
1.2	1.184	5.84	3.513

de Heen, 1881

%	b.t.
7.27	101.8
14.58	105.0
21.36	108.0

Gerlach, 1886

%	b.t.	%	b.t.
0	100	30.26	116
4.65	101	30.98	117
8.43	102	31.69	118
11.66	103	32.38	119
14.31	104	33.07	120
16.60	105	33.68	121
18.37	106	34.30	122
20.00	107	34.89	123
21.57	108	35.48	124
23.02	109	36.06	125
24.07	110	36.59	126
25.71	111	37.11	127
26.79	112	37.62	128
27.75	113	38.12	129
28.67	114	38.61	130
29.48	115		

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
3.26	100.42	14.39	103.20
5.84	100.86	16.95	104.30
8.39	101.37	18.07	104.77
12.18	102.40		

Gerlach, 1926

%	b.t.	%	b.t.
2.6	100.5	17.3	107.8
5.0	101.2	20.0	107.1
5.3	102.3	21.2	110.1
9.9	102.6	26.1	111.5
10.5	105.3	30.0	115.7
16.8	105.2	35.0	123.2

Van't Hoff and Meyerhoffer, 1898

mol %	f.t.	mol %	f.t.
0	0	9.28	-18.22
0.96	- 3.06	9.43	-19.4
2.03	- 7.65	9.90	- 3.4 E α
2.85	-13.65	9.93	- 9.6 E β
4.93	-33.6	9.99	0
6.41	-22.4	10.41 (12+1)	22
7.14	-18.8	11.57	62
8.33	-16.4	11.76	68.2
8.42	-16.45	12.35	79.7
8.49	-16.52	16.18	116.67 (4+1)
8.64	-16.60	18.18 (8+1)	152.6
8.75	-16.8	23.81	181-182 (2+1)
9.01	-17.4	24.10	186

Jones and Chambers, 1900; Jones, 1904
Jones and Getman, 1904; Jones and Bassett, 1905

M	f.t.	M	f.t.
0.0508	-0.280	0.6099	- 3.472
0.1016	-0.537	0.464	- 2.482
0.1525	-0.771	0.927	- 6.140
0.2033	-1.058	1.391	-10.820
0.2541	-1.335	1.854	-17.380
0.3801	-2.015	2.318	-27.000
0.5082	-2.762		

Menshutkin, 1907

%	f.t.	%	f.t.
40.18	86	42.90	107
40.70	92	45.33	115

Jones and Stine, 1908

M	%	f.t.	M	%	f.t.
0.15	1.412	-0.747	0.85	7.597	-5.404
.25	2.334	-1.306	1.05	9.262	-7.191
.35	3.241	-1.910	1.25	10.88	-9.236
.45	6.634	-2.537	1.45	12.46	-11.47
.65	5.899	-3.854			

Biltz and Marens, 1911

%	f.t.
34.49	3.5
35.59	25
36.97	50

The authors correct the value given by Van't Hoff and Meyerhoffer, 1898 for 25° (in interpolating their own measures for 0° and 62°) as 9.45 mol %.

Rodebush, 1918

%	f.t.	%	f.t.
8.08	-5.71	16.92	-19.57
10.50	-8.46	19.49	-25.86
14.02	-13.79	21.88	-33.50 E

Queisner, 1921

%	f.t.	%	f.t.
	(6+1)		
35.14	20	38.55	70
35.68	30	39.67	80
36.27	40	40.97	90
37.00	50	42.44	100
37.76	60		

Küpper, 1926

%	f.t.	%	f.t.
	(6+1)		
34.60	0	37.20	50
34.75	5	37.55	55
34.90	10	37.69	57
35.11	15.3	37.83	59
35.22	20	37.90	60
35.30	20.6	38.00	61
35.32	25	38.28	64
35.36	27	38.85	70
35.68	30	38.93	71
35.80	31	39.80	80
36.19	35.5	40.16	83
36.50	40	41.00	90
37.15	49	42.20	100

Akhumov and Vasiliev, 1932

%	f.t.
42.36	100 (6+1)
48.53	130 (6+1) - (4+1)
51.63	150 (4+1)
54.55	170 (4+1) - (2+1)
56.60	200 (2+1)
59.40	220
60.78	230
67.83	300

Chernogorenko, 1956

E_1 : -33.8 (6+1) E_2 : -21.9 (12+1)

Properties of phases .

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0154	18	1.1585
4	.0326	20	.1778
6	.0499	22	.1973
8	.0674	24	.2174
10	.0850	26	.2377
12	.1030	28	.2587
14	.1212	30	.2800
16	.1397		

Kremers, 1858

%	d	%	d
	19.5°		
0	0.9983	19.345	1.1683
7.732	1.0640	26.222	.2345
13.857	.1177	32.328	.2994

t	d				
19.5	1.0640	1.1176	1.1682	1.2346	1.2994
0	1.0677	1.1224	1.1737	1.2407	1.3061
40	.0573	.1106	.1610	.2271	.2918
60	.0486	.1021	.1527	.2191	.2838
80	.0381	.0922	.1434	.2103	.2757
100	.0261	.0812	.1331	.2008	.2666

Schiff, 1858 - 1859

%	d	%	d
	24°		
0	0.9973		
0.94	1.0042	19.66	1.1488
1.87	.0111	20.59	.1567
2.81	.0180	21.53	.1646
3.74	.0248	22.46	.1725
4.68	.0317	23.40	.1804
5.62	.0387	24.34	.1886
6.55	.0457	25.27	.1968
7.49	.0528	26.21	.2050
8.42	.0598	27.14	.2134
9.36	.0669	28.08	.2219
10.30	.0741	29.02	.2305
11.23	.0813	29.95	.2392
12.17	.0886	30.89	.2480
13.10	.0959	31.82	.2598
14.04	.1032	32.76	.2658
14.98	.1107	33.70	.2749
15.91	.1187	34.63	.2840
16.85	.1258	35.57	.2933
17.78	.1334	36.50	.3028
18.72	.1410	37.44	.3124

Gerlach, 1859

%	d	%	d
	15°		
0	0.99913	18	1.15821
1	1.00756	19	.16760
2	.01600	20	.17697
3	.02433	21	.18683
4	.03288	22	.19670
5	.04132	23	.20657
6	.05040	24	.21644
7	.05878	25	.22630
8	.06751	26	.23669
9	.07624	27	.24708
10	.08597	28	.25747
11	.09400	29	.26786
12	.10302	30	.27825
13	.11203	31	.28916
14	.12105	32	.30007
15	.13007	33	.31098
16	.13945	34	.32189
17	.14884	35	.33280

Schiff, 1861

%	d
15°	
0	0.9991
4.47	1.0364
8.94	1.0747
13.41	1.1141
17.88	1.1551
26.82	1.2445

Oudemans Jr., 1868

%	d	%	d
	14°		
0	0.9993	11.73	1.1018
0.47	1.0033	12.40	.1061
0.94	.0073	12.77	.1103
1.41	.0113	13.24	.1146
1.88	.0154	13.70	.1189
2.34	.0194	14.06	.1232
2.81	.0234	14.53	.1275
3.27	.0274	15.00	.1319
3.75	.0314	15.47	.1363
4.22	.0355	15.94	.1407
4.69	.0395	16.41	.1451
5.16	.0435	16.88	.1495
5.63	.0476	17.34	.1540
6.10	.0517	17.81	.1584
6.57	.0558	18.27	.1628
7.04	.0599	18.75	.1673
7.51	.0641	19.22	.1718
7.98	.0682	19.69	.1763
8.44	.0724	20.16	.1809
8.91	.0765	20.63	.1855
9.38	.0807	21.10	.1901
9.85	.0849	21.57	.1948
10.32	.0891	22.04	.1995
10.79	.0933	22.51	.2042
11.26	.0976		

Kohlrausch and Grotrian, 1875

%	d	%	d
	18°		
5	1.0416	25	1.2257
10	1.0861	30	1.2780
15	1.1295	34	1.3212
20	1.1765		

Kohlrausch, 1879

%	d	%	d
	18°		
5	1.0416	30	1.2779
10	1.0859	34	1.3210
20	1.1764		

Chernai, 1889

t	d	t	d
	5.01 %		
0.00	1.0425	26.37	1.0350
4.27	1.0422	34.72	1.0338
9.09	1.0416	38.10	1.0330
13.12	1.0403	40.27	1.0315
16.36	1.0398	43.97	1.0288
18.45	1.0391	50.58	1.0210
	17.43 %		
0.00	1.1527	25.14	1.1457
4.50	1.1517	30.30	1.1439
10.00	1.1502	34.18	1.1426
15.13	1.1488	40.43	1.1402
20.28	1.00416 (?)		

Charpy, 1893

%	d	%	d
	0°		
0	0.9999	15.7989	1.1426
6.2008	1.0544	20.9293	1.1925
11.3249	1.1006	29.2056	1.2786

Barnes and Scott, 1898

%	d	%	d
	20.1°		
0	0.9982	10.185	1.0833
1.743	1.0126	15.79	1.1324
3.903	1.0240	20.31	1.1735
3.913	1.0304	25.59	1.2241
5.919	1.0473	28.83	1.2569
8.058	1.0650		

Bremer, 1902

t	d	t	d
	20.0004 c		
14.92	1.156238	40.71	1.147256
15.06	.156220	50.26	.143319
20.45	.154419	59.67	.139316
27.76	.151958	69.82	.134639
33.72	.149856	100.23	.118401
	13.3111 c		
8.84	1.108366	41.99	1.097757
9.18	.108340	51.42	.093830
19.72	.105573	61.69	.089160
27.57	.102472	70.32	.085179
34.79	.100530	100.13	.068707
	9.9506 c		
16.44	1.08061	54.32	1.06701
21.55	.07939	65.28	.06182
30.59	.07649	73.62	.05726
37.74	.07386	82.04	.05261
47.43	.06999	100.19	.04202
	6.7158 c		
15.99	1.055595	55.84	1.041599
26.70	.052938	65.15	.036967
32.95	.050828	75.08	.031715
37.79	.049133	84.18	.026399
46.48	.045690	100.27	.016485

c = g in 100 cc at 15.36° (for the 6.7158 % it is stated at 11.76° ?)

Jones and Stine, 1908

M	%	d	M	%	d
		0°			
0.15	1.412	1.011816	0.85	7.597	1.065840
.25	2.334	.020288	1.05	9.262	.080044
.35	3.241	.028264	.25	10.88	.094152
.45	6.634	.035300	.45	12.46	.108380
.65	5.899	.049512			

Barnes and Johnson, 1902

c	d	c	d
	25.5°		
0	0.9969	10.85	1.0787
2.60	1.0164	12.00	.0867
4.24	.0290	13.50	.0997
5.36	.0375	14.70	.1077
7.70	.0552	16.70	.1200
8.88	.0640	21.19	.1518
9.70	.0701		

Grabowsky, 1904

%	d 10°	d 18°	d 30°
0	0.999731	0.99863	0.99567
7.96	1.0710	1.0679	1.0631
14.96	1.1318	1.1296	1.1261
21.46	1.1924	1.1903	1.1854

Cheneveau, 1907

%	d	%	d
13°		15°	
0	0.9994	0	0.9991
2.29	1.0175	3.69	1.0313
4.49	.0363	7.18	.0605
5.57	.0469	8.85	.0754
6.62	.0566	10.48	.0901
8.68	.0744	13.61	.1186
10.67	.0905	16.58	.1460

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	0.9804	1.03710
0.0993	1.00265	1.9538	.07355
.1987	.00652	4.3147	.15700
.4912	.01815		

Buchanan, 1912

m	d	m	d
	19.5°		
1.0	1.071615	5.0	1.292846
2.0	.134526	5.5	.315561
3.0	.191990	5.9182	.333868
4.0	.244582	5.9820	.336557

N	d	N	d
	19.5°		
0	0.998337	0.031	1.000785
0.001	.998410	.062	.003213
.002	.998500	.125	.007987
.004	.998626	.250	.017392
.008	.998963	.500	.035650
.016	.999570		

Herz, 1914-17

N	d
25°	
0	0.9971
1.40	1.0489
2.80	.0994
4.20	.1476
5.60	.1941

Pulvermacher, 1920

N	d
25°	
0.856	1.0300
1.712	.0630
3.424	.1234
6.848	.2367

Queisner, 1921

%	t	d	%	t	d
		sat.sol. (6+1)			
35.14	20	1.3368	37.76	60	1.3482
35.68	30	.3388	38.55	70	.3545
36.27	40	.3404	39.67	80	.3592
37.00	50	.3448	40.97	90	.3668
			42.44	100	.3727

Küpper, 1920

c	d	c	d
	20°		
5.16	1.0410	30.06	1.2224
10.74	1.0839	46.68	1.3324
19.96	1.1510		

Koch, 1924

t	d	t	d
	7.40 %		
- 1.9	1.0648	13.5	1.0625
- 1.4	1:0647	18.4	1.0612
- 0.2	1.0646	19.0	1.0611
+ 3.4	1.0642	19.5	1.0609
+ 8.1	1.0635		
	14.49 %		
-12.3	1.1310	8.1	1.1270
- 8.3	1.1305	15.3	1.1253
- 0.5	1.1291	18.5	1.1243
+ 1.7	1.1287		
	22.11 %		
-25.7	1.2074	+ 6.4	1.1998
-20.8	1.2064	16.1	1.1969
- 9.7	1.2040	18.9	1.1960
- 1.4	1.2019		
	28.71 %		
-14.5	1.2734	7.4	1.2668
-13.8	1:2732	15.0	1.2640
- 8.5	1.2716	18.7	1.2630
+ 0.3	1.2690		

Manchot, Jahrstorfer and Zepter, 1924

%	d
25°	
10.511	1.0802
20.165	1.1501

Cabrera and Duperier, 1925

28.3 % 20° d = 1.2516

Gerlach, 1926

%	d	%	d
	at room t.		
2.6	1.022	17.3	1.151
5.0	1.041	20.0	1.177
5.3	1.044	21.2	1.189
9.9	1.084	24.6	1.222
10.0	1.085	26.1	1.238
10.5	1.090	30.0	1.278
16.8	1.147	35.0	1.332

Herz and Hiebenthal, 1929

N	d 25°	d 70°
0.5	1.0159	0.9983
1.0	.0349	1.0173
2.0	.0716	.0542
4.0	.1409	.1244

Volkmann, 1930

%	d 15-16°
0	0.999
4.39	1.0362
9.87	.0840
19.21	.1694
25.71	.2338

Akhumov and Vasiliev, 1932

%	t	d
42.36	100	1.382
48.53	130	.411
51.63	150	.428
54.55	170	.442
56.60	200	.462

Spacu and Popper, 1934

20°

%	d	%	d
0	0.99823	27.566	1.251340
19.043	1.166740	31.700	.295391
19.630	.172360	31.855	.296880
23.785	.212830		

Okazaki, 1935

28°

%	d	%	d
6.07	1.0459	23.10	1.1975
10.91	1.0870	27.88	1.2446
16.76	1.1384		

Rutskov, 1948

mol%	d 25°	d 50°	d 75°
0.1	1.00145	0.9924	0.9793
0.25	.0078	.9988	.98575
7.7	.2731	1.2634	1.2530
9.0	.3139	.3043	.2940
9.46	.3318	-	-
10	-	1.3308	1.3208

De Heen, 1881

t	relative vol.	t	relative vol.	t	relative vol.
21.36%		14.58%		7.27%	
10.00	1.000000	10.00	1.000000	10.00	1.000000
15.28	.001375	15.30	.001233	15.40	.001125
20.62	.002842	20.32	.002538	21.12	.002466
28.05	.005058	28.34	.004874	28.77	.006903
35.38	.007296	35.28	.007100	43.15	.009648
41.22	.009299	41.12	.009103		
49.90	.012309	48.58	.011871		
57.48	.015158	56.65	.015080		
67.56	.019165	65.02	.018715		
		73.08	.022365		

Grotrian, 1876

%	$\eta^{18°}$	$\tau.10^4$
4.51	1704	318
19.83	4549	336
33.6	27460	406

Herz, 1914 and 1917

25°

N	η	N	η
0	895	4.200	2183
1.400	1172	5.600	3032
2.800	1593		

Pulvermacher, 1920

N	η(water=1) 25°
0.856	1.187
1.712	1.398
3.424	1.993
6.848	4.604

Herz and Hiebenthal, 1929

N	η 25°	η 70°
0.5	972.8	448.1
1.0	1078.0	489.6
2.0	1310.5	592.1
4.0	2010.6	851.2

Stakelbeck and Plank, 1929

%	η +20°	η +10°	η 0°	η -10°
0.9	1030	1390	1830	-
1.7	1060	1430	1910	-
2.6	1090	1470	1990	-
3.5	1130	1520	2070	-
4.4	1170	1570	2160	-
5.2	1210	1620	2250	-
6.1	1260	1680	2350	-
7.0	1310	1740	2450	-
7.1	1370	1820	2570	-
8.7	1430	1900	2670	-
9.6	1490	1990	2820	-
10.5	1550	2080	2950	-
11.4	1620	2180	3100	4600
12.3	1700	2310	3270	4800
13.2	1790	2440	3440	5000
14.1	1900	2600	3640	5250
15.1	2020	2760	3850	5550
16.1	2150	2940	4080	5950
17.1	2290	3120	4320	6400
18.1	2440	3300	4600	7500
19.1	2600	3540	4920	8200
20.1	2810	3990	5570	8600
21.1	3030	4150	5830	8900
22.1	3270	4490	6420	9800
23.1	3560	4860	7090	10900
24.1	3870	5320	7820	12150
25.1	4200	5830	8660	13630
26.2	4600	6400	9590	15180
27.2	5010	7040	10550	17130
28.3	5510	7760	11680	19120
29.4	6190	8550	12840	21250
30.4	6890	9200	14050	23900
31.5	7590	10000	15300	26400
32.7	8450	11120	16720	29400
33.9	9390	12310	18210	-
	-15°	-20°	-25°	
14.1	6900	-	-	
15.1	7280	-	-	
16.1	7580	9350	-	
17.1	8130	10020	-	
18.1	8600	10700	12900	
19.1	9190	11600	14200	
20.1	10020	12600	15790	
20.6	10500	13000	16500	
21.1	11030	13800	17430	
22.1	12280	15200	19430	
23.1	13790	17000	22800	
24.1	15480	19200	25950	
25.1	17600	22000	-	
26.2	20020	25200	-	
27.2	22420	-	-	
28.3	24950	-	-	

Ochsé, 1890

t	a^2 0 c	5 c	10 c	15 c	20 c
4	77.518	65.856	64.484*	61.348	58.898
8	75.362	-	-	-	-
15	72.912	62.622	61.250	59.290	57.624
25	70.168	59.780	58.996	56.938	56.546
35	64.190	57.330	56.840	54.488	55.762
45	60.270	54.880	54.488	52.430	54.586

*10 c 64.484 at 6°

Grabowsky, 1904

%	σ 10°	σ 30°
0	74.01	71.03
7.96	77.00	74.10
14.96	79.92	77.00
21.46	83.30	80.31

Volkmann, 1930

%	σ	%	σ
	15-16°		
0	73.2	19.21	82.0
4.39	75.2	25.71	86.7
9.87	77.2		

Kremers, 1859

d	t	n_D
1.1206	15	1.3683
1.2076	17	1.3936
1.3386	16	1.4311

Cheneveau, 1907

%	n_D	%	n_D
13°		15°	
0	1.3335	0	1.3334
2.29	.3391	3.69	.3426
4.49	.3448	7.18	.3515
5.57	.3475	8.85	.3559
6.62	.3502	10.48	.3602
8.68	.3557	13.61	.3686
10.67	.3608	16.58	.3768

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.9804	1.34488
0.0993	.33447	1.9538	.35585
0.1987	.33565	4.3147	.38064
0.4912	.33920		

Pulvermacher, 1920

N	n_D	N	n_D
	25°		
0.856	1.3434	3.424	1.3714
1.712	.3531	6.848	.4044

Spacu and Popper, 1934

%	n_D	%	n_D
	20°		
0	1.3324865	27.566	1.406842
19.043	.382365	31.700	.419484
19.630	.383957	31.855	.419878
23.785	.395728		

Akhumov and Golovkov, 1935

%	n_D	%	n_D
	25°		
0	1.3334	27.75	1.4082
7.65	.3521	35.22	.4295
14.00	.3690	46.86	.510
20.40	.3866	100	.6325

Kohlrausch and Grotrian, 1875

%	κ 0°	κ 18°
5	429	679
10	706	1114
15	-	1229
20	-	1394
25	-	1297
30	606	1055
34	-	758

Grotrian, 1876

%	κ	$\tau.10^4$
	18°	
4.51	622	217
19.83	1394	230
33.60	789	306

Kohlrausch, 1879

%	κ	$\tau.10^4$
	18°	
5	676	223
10	1120	221
20	1393	238
30	1054	284
34	762	319

Barnes and Johnson, 1902

c	κ	c	κ
	25.5°		
2.60	2.359	10.85	7.425
4.24	3.715	12.00	7.903
5.36	4.274	13.50	8.307
7.70	5.691	14.70	8.464
8.88	6.441	16.70	8.807
9.70	6.870	21.19	9.854

Jones, 1904. Jones and Getman, 1904. Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.0508	88.64	0.6099	66.59
0.1016	83.30	0.464	64.39
0.1525	79.11	0.927	63.43
0.2033	76.28	1.391	52.65
0.2541	74.84	1.854	45.70
0.3801	70.51	2.318	36.67

Jones and Stine, 1908

M	%	κ
	0°	
0.15	1.412	136.42
.25	2.334	215.29
.35	3.241	282.4
.45	6.634	361.83
.65	5.899	478.69
.85	7.597	576.06
1.05	9.262	662.80
.25	10.88	734.99
.45	12.46	780.40

Olmer, 1938

% (6+1)	κ	%	κ
	17°		
0	very low	40	1615
5	455	45	1540
10	800	50	1430
15	1000	55	1280
20	1220	60	1205
25	1370	65	1065
30	1470	70	930
35	1585	75	835

Cabrera and Dupérier, 1925

28.3% 20° χ= -0.665

Okazaki, 1935

%	Verdet's constant.10^5 (3441 Å)
28°	
6.07	4927
10.91	5366
16.76	5931
23.10	6527
27.88	6985

Okazaki, 1942

%	Verdet's constant.10^5 (D)
20°	
10.05	1532
13.49	1620
17.59	1717
20.81	1798
24.62	1897
28.42	1996

Mc Clung and Mc Intosh, 1902

% of X-ray absorption	d
room temperature	
81.8	1.140
69.1	.071
64.3	.036
59.5	.017
56.9	.000

Thermal constants.

Marignac, 1876

%	U
22-52°	
26.11	0.6824
17.49	.7716
9.58	.8665
5.03	.9235
2.58	.9594

Gumlich and Wiebe, 1898

t	U d°=1.1632	U d°=1.1493
-12	0.746	0.766
-5	.747	.774
-2	.754	.775
0	.750	.781

Küpper, 1920

c	U
20-100°	
5.16	0.9253
10.74	.8556
19.96	.7955
30.06	.7058
46.68	.6008

Koch, 1922

%	U -30°	U -20°	U -10°
12	-	-	0.821
14	-	-	.794
16	-	0.762	.768
18	-	.739	.745
20	0.712	.719	.724
22	-	.698	.708
24	-	.678	.684
26	-	.659	.665
28	-	-	.647
30	-	-	.627
32	-	-	.609

%	0°	10°	20°	30°	40°
0	1.005	1.001	0.999	0.998	0.998
2	0.973	0.973	.974	.974	.974
4	.943	.944	.945	.947	.948
6	.913	.915	.915	.920	.923
8	.883	.886	.889	.893	.896
10	.853	.857	.861	.865	.870
12	.825	.830	.835	.840	.845
14	.799	.804	.810	.815	.821
16	.774	.780	.786	.792	.799
18	.751	.758	.765	.771	.777
20	.730	.737	.743	.749	.756
22	.710	.716	.723	.729	.736
24	.690	.697	.703	.710	.716
26	.671	.677	.684	.690	.697
28	.653	.659	.665	.671	.678
30	.633	.640	.646	.652	.659
32	.615	.621	.627	.634	.640
34	.596	.602	.606	.615	.621
66	-	-	-	-	.602

Gerlach, 1926

%	U	t
2.6	0.9581	18-52
5.0	.9253	20-100
5.3	.9245	-
9.9	.8556	20-100
10.0	.8530	-
10.5	.8607	-
16.8	.7660	- 8
17.3	.7810	- 1
20.0	.7120	-30
20.0	.7300	0
20.0	.7560	40
21.2	.7716	-
24.6	.7058	20-100
26.1	.6824	22-52
30.0	.6330	0
30.0	.6590	40
35.0	.6008	20-100

Nikolaev, Kogan and Ogorodnikov, 1936

%	U
	25°
2.00	0.9680
14.76	.7943
18.05	.7543
21.64	.7188
26.00	.6802
32.30	.6234

Rutskov, 1948

%	U		
	25°	50°	75°
0.526	0.9900	0.9904	-
1.00	.9821	-	-
1.304	.9779	0.9787	0.9828
2.575	.9586	.9601	.9650
5.021	.9223	.9248	.9310
9.562	.8616	.8654	.8721
14.98	.7968	.8012	-
20.91	.7330	.7387	.7453
24.84	.6928	.7002	-
30.58	.6388	.6452	.6528
34.58	.6070	.6135	.6215
35.60	.5992	-	-
37.00	-	.5970	.6053

Jäger, 1891

%	relative heat conductivity coefficient
0	100
11	94.9
14.5	91.7
22	89.0
29	85.4

Kapustinskii and Ruzavin, 1955

%	K	%	K
	25°		
4.70	1439	20.0	1342
9.70	1413	28.3	1288
14.6	1380		

heat conductivity coefficient.10^6

Water + Magnesium bromide ($MgBr_2$)

Tammann, 1885

t	p 0%	14.15%	21.40%	28.10%	36.73%
24.31	22.8	21.2	20.2	18.2	15.2
28.27	28.8	27.3	25.5	23.1	19.5
32.46	36.6	34.5	32.5	29.5	24.4
36.36	45.4	42.7	40.4	36.6	30.6
39.71	54.5	51.3	48.7	44.5	36.9
44.23	69.1	65.2	62.0	56.3	47.0
49.03	88.2	83.7	79.3	72.5	60.4
52.90	106.7	100.8	95.5	87.4	72.9
57.82	135.0	127.5	121.4	110.8	93.0
59.51	146.1	139.3	131.8	120.8	101.3
63.40	174.6	168.4	157.5	144.0	120.8
65.73	193.9	185.5	175.8	160.8	135.3
68.97	223.6	214.0	203.3	185.2	156.0
72.47	260.0	248.5	235.2	219.4	182.4
74.76	286.4	271.8	258.5	241.3	200.8
76.90	313.1	298.0	282.7	263.3	219.6
78.96	340.9	324.6	308.7	286.9	239.6
81.38	375.8	357.1	339.4	314.5	263.9
85.07	435.1	411.9	393.0	361.3	306.9
87.34	475.1	452.4	429.9	397.6	335.4
90.66	539.8	512.4	487.5	451.4	332.4
93.23	593.8	564.6	537.2	497.2	421.2
96.57	671.5	640.1	609.6	562.3	478.2
98.44	718.7	682.6	649.8	602.1	511.2
99.95	758.7	720.8	686.8	635.8	541.1

100°

%	p	%	p
6.69	746.5	39.94	491.7
11.09	734.4	40.72	484.4
16.63	712.6	46.17	397.0
24.36	666.8	46.54	366.9
31.23	605.6		

Lescoeur, 1894

t	p sat.sol.	t (1+1)	p dissoc.
20	3.4	120	-
60	45	130	30
80	89	140	51
100	167	150	81
117	273	160	146
		165	205

Robinson and Stokes, 1940

m_1	m_2	m_1	m_2
25°			
0.1086	0.1557	0.1127	0.1614
.3250	.5101	.3575	.5668
.5171	.8734	.6057	1.060
.8688	1.659	1.086	2.307
1.576	3.697	.667	4.009
.802	4.462	.894	4.782
0.1341	0.1939	0.1770	0.2600
.4284	.6988	.4592	.7607
.6845	1.240	.7192	1.303
1.304	2.852	1.556	3.627
.697	4.092	.761	4.301

m_1 = molality of magnesium bromide

m_2 = molality of potassium chloride

Stokes, 1948

m	osmotic coefficient
25°	
0.1	0.874
.2	.898
.3	.928
.4	.963
.5	1.004
.6	.042
.7	.082
.8	.127
.9	.172
1.0	.218
1.2	.314
1.4	.410
1.6	.510
1.8	.610
2.0	.715
2.5	.999
3.0	2.29
3.5	.59
4.0	.89
4.5	3.19
5.0	.50
5.84	.85

Etard, 1894

%	f.t.
52.0	-18
58.0	+17
60.9	48
62.5	62
65.8	97

Jones, 1904. Jones and Getman, 1904. Jones and Bassett, 1905

M	f.t.	M	f.t.
	0°		
0.0517	-0.277	0.517	-3.022
.103	-0.531	.642	-3.921
.155	-0.801	.964	-6.850
.207	-1.088	1.610	-15.200
.310	-1.690	2.571	-37.500
.321	-1.691		
.414	-2.347		

Menshutkin, 1907

%	f.t.	%	f.t.
47.87	0	56.07	120
49.20	20	57.96	140
50.52	40	59.79	150
51.84	60	61.73	160
53.16	80	62.99	164
54.55	100		

Getman, 1935

%	f.t.	%	f.t.
0	0	49.32	+0.83 (tr.t.)
10.71	-3.92	49.34	10.4
15.43	-6.85	49.83	19.9
24.61	-15.2	50.23	24.8
35.06	-37.5	50.78	29.8
36.75	-42.7	51.33	34.8
40.63	-23.0	51.72	39.8
42.38	-15.0	52.64	60.1
44.00	-9.0	53.36	65.5
44.20	-8.0	55.61	100.0
44.32	-7.0	62.99	172.4
44.83	-6.0		
49.54	-0.7		
(10+1)	(6+1)		

Denecke, 1919

P Kg	f.t.	P Kg	f.t.
1	-43.3	1483	-54.6
601	47.9	2596	58.4
1222	52.3	2915	60.5
E	38% (10+1) + ice I		

Kremers, 1858 and 1859

%	d	%	d
	19.5°		
5	1.041	30	1.308
10	.085	35	.375
15	.135	40	.449
20	.189	45	.532
25	.245	50	.622

t	d				
19.5	1.1542	1.2826	1.3751	1.4723	1.5800
0	1.1593	1.2897	1.3832	1.4817	1.5896
40	.1463	.2736	.3654	.4620	.5693
60	.1368	.2635	.3550	.4514	.5596
80	.1258	.2523	.3438	.4401	.5469
100	.1129	.2400	.3318	.4283	.5358

Jones, 1904. Jones and Getman, 1904. Jones and Bassett, 1905

M	d	M	d
	0°		
0.0517	1.006104	0.414	1.077844
.103	.014160	.517	.093964
.155	.023960	.642	.139940
.207	.030212	1.610	.231720
.310	.044340	2.571	.366956
.321	.063176	3.214	.446324

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.0025	1.0	1.0739
.10	.0064	2.0	.1480
.20	.0140	4.0	.2922
.50	.0369		

Kremers, 1859

d	n_D
17°	
0.9983	1.3332
1.5006	.4322
1.2313	.9793

Jones, 1904. Jones and Getman, 1904. Jones and Bassett, 1905

M	n_D	M	n_D
	0°		
0.321	1.33518	1.610	1.37350
.642	.34536	2.571	.40034
.964	.35463	3.214	.41816

Jones, 1904. Jones and Getman, 1904. Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.0517	100.30	0.517	77.04
.103	94.55	.642	75.56
.155	90.62	.964	68.31
.207	88.49	1.610	54.63
.310	82.66	2.571	37.28
.321	81.86		
.414	79.90		

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	94.5	1.0	65.2
.10	88.8	2.0	54.6
.20	83.1	4.0	38.1
.50	73.2		

Jauch, 1921

N	U
18°	
0.5	0.9441
1	.8940
2	.8068
3	.7333
4	.6688

Water + Magnesium iodide (MgI_2)

Robinson and Stokes, 1940

m_1	m_2	m_1	m_2
	25°		
0.1049	0.1518	0.1217	0.1789
0.2696	0.4233	.3595	.5934
0.6263	1.153	.6363	1.178
0.8704	1.776	.9268	.933
1.218	2.803	1.219	2.821
1.621	4.188	.688	4.423
0.1266	0.1867	0.1958	0.2982
.4989	.8626	.6225	1.152
.7496	1.449	.7598	.484
.9707	2.058	1.162	2.621
1.440	3.543	.515	3.806
.743	4.647	.787	4.81

m_1 = m of MgI_2 m_2 = m of KCl

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.892	1.4	1.537
0.2	0.921	1.6	1.660
0.3	0.957	1.8	1.784
0.4	0.998	2.0	1.912
0.5	1.044	2.5	2.25
0.6	1.090	3.0	2.60
0.7	1.139	3.5	2.96
0.8	1.192	4.0	3.34
0.9	1.249	4.5	3.72
1.0	1.306	5.0	4.11
1.2	1.421		

Menshutkin, 1907

%	f.t.	%	f.t.
(8+1)		(6+1)	
54.71	0	64.65	43
58.32	20	65.01	80
63.36	40	65.45	120
65.78	43.5	66.02	160
		67.25	200
		67.89	215

Kremers, 1859

%	$d^{19.5°}$	n_D 16°
-	0.9983	1.3333
-	1.3234	.3975
-	1.6305	.4594
60.22	1.9065	.5198

Kremers, 1860

%	d	%	d
	19.5°		
0	0.998	35	1.393
5	1.041	40	.472
10	.086	45	.565
15	.137	50	.665
20	.192	55	.777
25	.252	60	.912
30	.318		

t			d		
19.5	1.19620	1.40220	1.52940	1.71630	1.83100
0	1.20226	1.41194	1.54077	1.72918	-
40	.18715	.39043	.51637	.70205	1.81641
60	.17623	.37767	.50277	.68762	.80167
80	.16365	.36392	.48842	.67267	.78665
100	.14964	.34933	.47349	.65752	.77158

Heydweiller, 1909

N	d	λ	N	d	λ
		18°			
0.05	1.0044	946	1.0	1.1122	686
.10	.0101	899	2.0	.2244	586
.20	.0215	837	4.0	.4472	410
.50	.0555	757			

Water + Magnesium fluosilicate (MgF_6Si)

Jatlov and Pinaevskaya, 1938

%	f.t.		%	f.t.
0	0		23.53	20
15.0	-0.9		25.86	40
10.0	-2.2		28.54	50
15.0	-3.8		28.66	57
19.5	-6.0	E	30.74	60
20.85	0	(6+1)		

Simpson and Glocker, 1953

% (6+1)	f.t.	% (6+1)	f.t.
34.83	1.7	45.31	43.3
37.37	10.0	46.94	51.7
39.71	18.3	48.46	60.0
41.76	26.7	49.82	68.3
43.60	35.0		

Water + Magnesium nitrite (MgN_2O_4)

Bureau, 1936 - 1937

%	f.t.	E	tr.t.	
7.04	-3.25	-21.15	-	
17.25	-14.3	"	-	
23.2	-21.15	"	-	
24.4	-16.15	"	-	E
26.6	-12.15	-21.8	-	(9+1)
30.25	-2.15	-21.15	-	
31.75	+1.15	-21.3	-	
38.60	10.5	"	10.5	
40.9	16.35	-20.0	"	(6+1)
47.0	25.65	-10.5	"	
52.0	29.5	"	29.5	
60.0	45	-	-	(3+1)

Water + Magnesium chlorate ($MgCl_2O_6$)

Meusser, 1902

%		f.t.	%		f.t.
22.24		- 8	63.82	(4+1)	42
26.35		-12	69.12		65.5
51.64	(6+1)	-18	65.37	(2+1)	39.5
53.27		0	69.46		61
56.50		18	70.69		68
60.23		29	73.71		93
63.65		35			

sat.sol. 18° d = 1.564

Rubien, 1911

N	d	n_D	N	d	n_D
		18°			
0	0.99862	1.33327	0.4963	1.03293	1.34522
0.1017	1.00579	.33574	1.0037	.13370	.35646
0.2025	.013955	.33927	3.9845	.26165	.37752

Heydweiller, 1912

N	d	ϰ	N	d	ϰ
		18°			
0.1012	1.00579	77.0	1.011	1.06805	554.5
0.2023	.01286	142.9	2.023	.13556	896
0.5057	.03390	317.6	4.046	.26560	1139

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34518
0.1	.33452	2.0	.35649
0.2	.33571	4.0	.33770
0.5	.33831		

Water + Magnesium bromate ($MgBr_2O_6$)

Linke, 1955

%	b.t.
15.15	101.0
27.2	102.6
44.3	106.5
60.9	116
72.2	131

%	f.t.	d
	sat.sol.	
9.34	-1.6	-
20.94	-4.2	-
34.66	-10.1	-
42.34	0 (6+1)	1.512
45.58	10	.562
48.66	20	.609
51.4	30	.662
54.5	40	.722
57.8	50	.787
62.5	65	.900
66.8	75	2.013
68.3	78	.070
70.15	80.5	-
	(2+1)	
70.1	81	-
70.8	90	-
71.9	100	-
72.6	117	-
73.3	130	-

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.0574	52.8
1	.1141	44.75
2	.2260	34
3	.3361	25.67

Water + Magnesium iodate (MgI_2O_6)

Mylius and Funk, 1897

%	f.t.	%	f.t.
(10+1)		(4+1)	
3.1	0	6.4	0
10.2	20	6.8	10
17.4	30	7.7	20
21.9	35	8.9	35
67.5	50	12.6	63
		19.3	100

Hill and Moskowitz, 1931

%	f.t.	d_t
	(10+1)	
3.18	- 0.36	1.026
4.39	+ 5	1.034
5.87	10	1.049
7.79	15	1.073
	anh.	
13.2	60	-
13.3	70	-
13.4	80	-
13.5	90	-
	(4+1)	
6.09	5	1.047
6.68	10	1.060
7.29	15	1.065
8.55	25	1.075
9.83	35	1.086
10.51	40	1.090
12.05	50	1.115
13.1	57.5	-
15.7	70	-
19.6	90	-

Water + Magnesium nitrate (MgN_2O_6)

Heterogeneous equilibria .

Tammann, 1885

%	p	%	p
	100°		
4.81	748.6	23.50	654.3
11.34	725.7	28.44	609.7
17.21	696.2	33.60	555.3

Sieverts and Petzold, 1932

mol%	p	mol%	p
30°		p dissoc.	
10.63	16.4	15.22	11.4
11.25	"	17.27	11.2
12.29	"	20.20	11.4
14.10	"	21.00	0.5
		24.33	"
		27.86	"
		31.25	"
		41.49	0.3

Ewing, Klinger and Brandner, 1934

%	p 20°	30°	40°	50°	60°
11.45	16.46	30.08	52.69	88.40	142.5
19.75	15.40	28.29	49.35	82.97	133.8
25.28	14.46	26.50	46.25	77.16	125.3
28.82	13.60	24.90	43.70	73.25	118.3
32.89	12.50	22.79	39.48	67.32	109.4
37.79	10.87	19.94	35.00	58.86	95.8
41.94	-	16.71	29.48	49.95	81.6
43.04	-	-	28.18	47.76	78.22
45.05	-	-	25.45	42.88	70.35
46.64	-	-	-	40.96	65.05
	25°	35°	45°	47°	53°
41.94	12.45	-	-	-	-
43.04	-	21.38	-	-	-
45.05	-	-	33.22	36.85	-
46.64	-	-	-	-	45.95
	55°	57°	70°		
43.04	-	-	124.3		
45.05	55.28	-	112.4		
46.64	50.89	56.28	103.7		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.857	.9	.046
.2	.869	1.0	.074
.3	.890	1.2	.134
.4	.914	1.4	.192
.5	.940	1.6	.253
.6	.967	1.8	.315
.7	.991	2.0	.380
.8	1.017		

Rüdorff, 1872

%	f.t.	%	f.t.
5.02	-1.85	12.12	-5.70
6.24	-2.40	14.97	-7.75
9.01	-3.70		

Funk, 1899

%	f.t.	%	f.t.
25.92	-17	42.33	+18
29.13	-22	45.87	40
32.60	-26	53.69	80
35.44(9+1)	-23	57.81	90
36.19	-20.5	63.14	89
38.03	-18	65.67	77.5
39.50(6+1)	-4.5	67.55	67
39.96	0		

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
	0°		
0.077	-0.370	1.236	- 8.650
.155	-0.753	.545	-12.000
.309	-1.552	.854	-16.270
.618	-3.559	2.163	-22.500
.927	-5.930		

Jones and Stine, 1908

M	f.t.	%.
0.15	-0.749	2.189
.25	-1.252	3.611
.35	-1.805	5.000
.45	-2.395	6.379
.55	-3.019	7.709
.65	-3.669	9.014
.75	-4.393	10.300
.85	-5.113	11.590
1.05	-6.762	14.060
.25	-8.562	16.380
.45	-10.589	18.780

Di Capua, 1929

43.68% f.t.=20°

Sieverts and Petzold, 1932

%	f.t.		
0	0		
10.6	-4.2		
16.6	-8.4		
16.9	-8.5		
20.8	-12.6		
24.6	-17.2		
26.1	-19.2		
30.2	-26.9		
30.6	-27.1		
32.6	-32.1		
33.4	-34.1		
32.3	-31.6	E	
32.0	-31.9	(9+1)	
32.4	-30.8		
33.0	-29.0		
33.3	-27.0		
34.3	-23.0		
36.1	-18.0		
37.2	-15.0		
35.3	-33.5	(6+1)	
35.9	-22.0		
37.5	-7.8		
38.5	+1.0		
40.4	15.0		
42.1	25.0		
43.2	35.0		
44.1	40.0		
45.8	50.0		
47.7	60.0		
50.3	75.0		
51.5	80.0		
53.3	85.0		
55.2	88.0		
57.0	89.0		
60.1	88.0		
61.6	85.0		
64.3	75.0		
65.7	65.0		
67.4	55.6	tr.t.	(6+1)-(2+1)
67.9	52.2	"	(6+1)-(4+1)
-	51.9	"	(4+1)-(2+1)
67.2	55.0	(2+1)	
67.9	65.0		
69.0	75.0		
70.7	90.0		
72.0	100.0		
73.3	110.0		
75.3	120.0		
77.2	124.5		
79.0	128.5		
81.4	128.5		
81.8	127.7	tr.t.	(2+1)-0aq.
82.0	130.0		
82.2	133.5		
82.2	136.0		
83.8	160.0		
84.6	186.0		

Ewing, Brandner, Slichter and Griesinger, 1933

%	f.t.	%	f.t.	%	f.t.
0.50	-0.232	39.6	+2.7	66.8	62.3
1.96	-0.476	40.4	10.2	67.2	56.9
2.87	-1.29	42.5	25.0	69.0	40.4
6.58	-2.49	43.2	32.4	68.7	52.9
12.09	-5.07	45.0	43.1	68.4	61.1
15.04	-6.84	46.7	49.3	69.3	65.9
18.80	-10.09	47.8	56.2	69.9	68.6
22.20	-13.82	48.9	60.0	70.0	71.8
25.50	-17.88	49.2	63.3	71.4	82.3
27.50	-20.17	50.2	66.8	71.4	85.4
29.30	-22.58	51.5	74.1	73.0	96.5
31.06	-25.66	52.4	77.6	74.4	114.4
32.37	-31.87 (9+1)	54.0	84.3	76.1	119.7
32.80	-30.50	55.8	87.8	77.5	125.8
33.60	-25.00	57.8	89.9	79.8	130.0
36.60	-15.30	61.2	87.6	81.1	130.9
37.10	-14.70E	64.2	81.2	82.4	130.0
36.70	-23.70	64.7	74.8	81.9	130.5
38.40	-4.80 (6+1)	66.4	64.5	82.1	137.5
				82.4,	144.6
					(2+1)

Vasiliev, 1916 and 1936

(2+1) f.t.= 127 E (6+1)+(2+1) f.t.= 55

Properties of phases .

Schiff, 1858 - 1859

%	d	%	d
	21°		
0	0.9980	15.03	1.1099
1.16	1.0059	16.18	.1195
2.31	.0139	17.35	.1291
3.47	.0210	18.51	.1388
4.62	.0301	19.66	.1486
5.78	.0385	21.82	.1586
6.94	.0470	21.97	.1687
8.09	.0557	23.12	.1789
9.25	.0643	24.28	.1892
10.40	.0732	25.43	.1997
11.56	.0822	26.58	.2102
12.72	.0913	27.74	.2209
13.87	.1005	28.90	.2318

Oudemans Jr., 1868

%	d	%	d
	14°		
0	0.9993	15.03	1.1151
0.58	1.0034	15.60	.1200
1.16	.0076	16.18	.1248
1.73	.0118	16.76	.1297
2.31	.0160	17.35	.1347
2.89	.0202	17.93	.1396
3.47	.0245	18.51	.1446
4.01	.0288	19.08	.1497
4.62	.0331	19.66	.1547
5.20	.0374	20.24	.1649
5.78	.0418	20.82	.1700
6.36	.0462	21.39	.1752
6.94	.0506	21.97	.1804
7.59	.0550	22.55	.1857
8.09	.0594	23.13	.1909
8.67	.0639	23.71	.1961
9.25	.0685	24.29	.2014
9.62	.0730	24.86	.2068
10.40	.0776	25.44	.2122
10.98	.0822	26.02	.2176
11.56	.0915	26.60	.2231
12.14	.0961	27.17	.2286
12.72	.1008	27.75	.2341
13.29	.1056	28.33	.2397
13.87	.1103	28.91	.2453
14.45	.0869		

Kohlrausch, 1879

%	d
	18°
5	1.0378
10	.0763
15	.1181
17	.1372

Chernai, 1889

t	d	
	7.59%	14.12%
0.00	1.0602	1.1156
4.17	.0595	.1144
10.00	.0583	.1124
14.93	.0570	.1105
20.11	.0554	.1085
26.27	.0533	.1059
30.39	.0517	.1040
35.78	.0495	.1014
40.41	.0475	.0991
44.95	.0454	.01502

Barnes and Scott, 1898

%	d	%	d
	20.1°		
0	0.9982	13.43	1.1028
1.372	1.0085	19.55	.1551
4.001	.0276	25.03	.2057
6.650	.0330	31.15	.2655
10.09	.0480	35.62	.3110

Jones, 1904; Jones and Getman, 1904

M	d	M	d
	0°		
0.077	1.007516	1.236	1.128436
.155	.015204	.545	1.161184
.309	.032520	.854	1.191968
.618	.065912	2.163	1.222468
.927	.098984	.472	1.253240

Chéneveau, 1907

%	d	%	d
	15°		
19.77	1.1517	7.52	1.0541
16.24	.1219	6.32	.0454
12.52	.0923	5.10	.0360
12.12	.0899	4.42	.0299
10.58	.0765	2.59	.0168
8.59	.0605	0	0.9991

Jones and Stine, 1908

M	%	d	M	%	d
		0°			
0.15	2.189	1.017224	0.75	10.30	1.080812
.25	3.611	.027520	.85	11.59	.087832
.35	5.000	.039128	1.05	14.06	.108320
.45	6.379	.048300	.25	16.38	.129740
.55	7.709	.058992	.45	18.78	.150504
.65	9.014	.070420			

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.00144	1.0	1.0528
.10	.00423	2.0	.1056
.20	.0098	4.0	.2009
.50	.0262		

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	1.016	1.05371
0.1028	1.00441	2.015	.10630
.2050	.01012	4.023	.20793
.5058	.02650		

Clausen, 1912

t	d			
	0.605 N	1.012 N	2.015 N	4.032 N
6	1.03392	1.05668	1.11065	1.21463
18	.03303	.05371	.10636	.20853
30	.03246	.04966	.10612	.20725

Herz, 1914

N	d
25°	
0	0.9971
1.56	1.0789
3.12	.1592
4.69	.2363
6.25	.3116

Herz, 1917

N	d
25°	
0	0.9971
1.562	1.0793
3.125	.1590
4.686	.2326
6.250	.3118

Manchot, Yahrstorfer and Zepter, 1924

%	d
25°	
14.389	1.0935
28.630	.1846
14.398	1.0916
28.184	.1821

Sieverts and Petzold, 1932

%	t	d
0	0	1.000
10.6	- 4.2	1.084
16.6	- 8.4	1.139
16.9	- 8.5	1.140
20.8	-12.6	1.181
24.6	-17.2	1.219
26.1	-19.2	1.233
30.2	-26.9	1.279
30.6	-27.1	1.281

Guillaume, 1946

%	d	%	d
	20°		
9.12	1.0711	23.07	1.1883
14.23	1.1138	39.75	1.3620
19.70	1.1624		

Wagner, 1883

%	η (water=1)			
	15°	25°	35°	45°
18.62	0.9979	0.8129	0.6653	0.5623
34.19	2.1332	1.6440	1.3242	1.0990
39.77	3.1705	2.5005	1.9140	1.5813

Herz, 1914

N	η (water=1)	N	η (water=1)
	25°		
0	1	4.69	2.470
1.56	1.290	6.25	3.685
3.12	1.747		

Herz, 1914 and 1917

N	η	N	η
	25°		
0	895	4.686	2211
1.562	1155	6.250	3298
3.125	1564		

Jones, 1904 and Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.077	1.32711	1.236	1.35222
.155	.32878	.545	.35870
.309	.33231	.854	.36468
.618	.33902	2.163	.37075
.927	.37579	.472	.37683

Chéneveau, 1907

%	n_D	%	n_D
	15°		
19.77	1.3639	7.52	1.3443
16.24	.3580	6.32	.3425
12.52	.3521	5.10	.3407
12.12	.3514	4.42	.3396
10.58	.3490	2.59	.3369
8.59	.3458	0	.3334

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34435
0.1	.33442	2.0	.35504
.2	.33554	4.0	.37510
.5	.33891		

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.016	1.34453
0.1028	.33446	2.015	.35521
.2050	.33560	4.023	.37534
.5058	.33898		

de Mallemann and Guillaume, 1945

20° d = 1.1582 (α) magn. . 10^5 = 13.40

Guillaume, 1946

%	(α)* magn . 10^6 5780 Å	n
	20°	
9.12	3.696	1.3482
14.23	.540	.3570
19.70	.370	.3668
23.07	.265	.3711
39.75	2.767	.4043

* in radians, gauss, centim.

Kohlrausch, 1879

%	κ	$\tau . 10^4$
	18°	
5	436	217
10	765	213
15	1015	209
17	1097	209

Jones, 1904 and Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.077	990.0	1.236	594.0
0.155	941.0	1.545	494.0
0.309	857.0	1.854	429.0
0.618	751.8	2.163	370.8
0.927	660.9	2.472	315.6

Jones and Stine, 1908

M	%	κ
	0°	
0.15	2.189	137.01
0.25	3.611	215.63
0.35	5.000	286.13
0.45	6.379	351.23
0.55	7.709	415.06
0.65	9.014	469.42
0.75	10.30	522.60
0.85	11.59	568.21
1.05	14.06	653.34
1.25	16.38	716.68
1.45	18.78	762.79

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	868	1.0	598
.1	820	2.0	488
.2	763	4.0	320
.5	681		

Clausen, 1912

t	λ	$\tau.10^5$
	0.605 N	
6	494.9	2996
18	658.0	3983
30	829.6	5022
	1.012 N	
6	454.7	4600
18	599.3	6063
30	754.3	7634
	2.015 N	
6	372.3	7500
18	487.0	9812
30	611.4	12320
	4.032 N	
6	243.6	9821
18	320.6	12920
30	402.1	16210

Marignac, 1876

%	U
21-52°	
3.96	0.9537
7.62	.9145
14.15	.8517
24.80	.7568
35.50	.6777

Water + Magnesium perchlorate ($MgCl_2O_8$)

Stokes, 1948

molality	osmotic coefficient
25°	
0.1	0.898
.2	.935
.3	.974
.4	1.016
.5	.062
.6	.108
.7	.158
.8	.211
.9	.267
1.0	.323
.2	.437
.4	.558
.6	.683
.8	.815
2.0	.945
2.5	2.306
3.0	.667
3.5	3.036
4.0	.397
4.4	.664

Copeland and Brag, 1954

mol%	$P \cdot 10^{-3}$	
23°		
0.50	0.56	(2+1)-0 aq.
.25	8.15	(4+1)-(2+1)
.166	20.9	(6+1)-(4+1)
sat.sol.	81	(6+1)

Lilich and Dzhurinskii, 1956

m	f.t.	m	f.t.
4.10	0	4.57	30
4.20	5	4.68	35
4.26	10	4.72	40
4.34	15	4.79	45
4.44	20	4.89	50
4.48	25		

Lilich and Mogilev, 1956

m	pH	m	pH
	25°		
0.0430	5.90	1.30	4.64
.392	.33	2.51	3.78
.667	.08	4.00	2.51

Water + Magnesium dithionate (MgO_6S_2)

De Baat, 1926

%	f.t.
31.94	0
33.30	12
33.91	20
35.24	30

Water + Magnesium sulfate (MgO_4S)

Heterogeneous equilibria .

Tammann, 1885

t	p 0%	p 8.8%	p 15.62%	p 16.31%	p 19.78%
45.11	72.3	71.0	69.8	69.4	68.8
54.84	117.2	115.7	113.7	113.4	111.8
59.26	144.4	142.5	140.5	139.8	138.5
63.83	177.2	174.6	171.9	171.7	169.2
67.67	211.2	208.6	204.9	204.4	202.1
71.18	246.0	242.9	239.0	238.7	235.8
57.84	299.7	296.0	292.2	291.7	287.7
79.37	346.5	342.5	337.4	336.8	336.1
83.56	409.8	404.9	399.8	398.5	393.9
86.13	453.3	448.6	442.3	440.6	436.5
89.86	523.2	517.7	-	508.2	504.2
92.42	576.2	569.5	562.0	560.7	-
95.65	649.4	640.5	633.9	632.4	626.9
100.37	770.0	760.3	751.4	750.0	743.2

%	p	%	p
100°			
9.44	749.2	21.48	729.6
14.93	743.3	28.71	704.2
20.21	733.6	33.66	677.3

Wiedemann, 1874

t	p dissoc.	t	p dissoc.
(x+1)			
11.6	9.6	50.0	75.7
22.2	16.0	51.9	81.9
24.3	17.8	60.0	122.5
29.8	26.5	61.2	129.2
30	28.0	61.8	132.5
35	35.6	70.2	188.9
40	47.2	70.4	190.3
40.2	46.3	80.0	276.0
45.4	60.9		

Frowein, 1887

p dissoc.	t	p dissoc.	t
(7+1) - (6+1)			
4.928	14.95	18.345	30.75
7.646	20.05	18.846	31.05

Lescoeur, 1890

t	p (x+1)	p (7+1)	p (6+1)	p sat.sol.
10	7.6	3.4	-	8
20	14.6	7.5	2.5	15.9
30	-	15.9	4.5	28
40	-	-	15	47.2
50	-	-	49	75.7
60	-	-	89	122.5
70	-	-	158	190.3
80	-	-	-	276

Cumming, 1909

p dissoc. (7+1) 20° 7.4 mm

Bolt, 1912

p dissoc.	hydrate	p dissoc.	hydrate
30.75°			
18.92	(7+1)	7.18	(4+1)
15.93	(6+1)	2.73	(2+1)
12.52	(5+1)		

Cohen, 1924

t	p dissoc.	t	p dissoc.
(7+1)			
30.75	18.96	43.40	48.37
40.02	37.90	47.45	64.31
41.42	41.89		

Pohle, 1927

t	p	t	p
sat.sol.			
30	30.0	70	187.0
35	38.0	75	232.0
40	48.0	80	285.0
45	61.0	85	343.0
50	78.0	90	408.0
55	99.5	95	492.0
60	125.0	100	587.0
65	154.0	105	693.0

Robinson and Stokes, 1949

m	osmotic coefficient
25°	
0.1	0.606
.2	.562
.3	.540
.4	.529
.5	.522
.6	.518
.7	.517
.8	.518
.9	.520
1.0	.525
1.2	.542
1.4	.567
1.6	.597
1.8	.630
2.0	.666
2.5	.780
3.0	.922

De Heen, 1881

%	b.t.
10.56	101.2
15.95	101.60
21.50	102.60

Gerlach, 1886

%	b.t.	%	b.t.
0	100	28.31	103
8.09	100.5	30.45	103.5
14.31	101	31.29	104
19.03	101.5	33.90	104.5
22.78	102	35.31	105
25.76	102.5	42.86	108

Kahlenberg, 1901

%	b.t.	%	b.t.
760 mm			
2.66	100.19	34.55	102.03
6.74	100.29	37.70	103.29
21.62	100.54	39.17	103.39
26.96	100.95	41.96	103.71
30.30	101.49		

Johnston, 1906

%	b.t.	%	b.t.
1.03	100.038	9.73	100.352
2.42	100.082	11.48	100.406
4.35	100.138	13.15	100.440
5.99	100.190	15.49	100.544
7.86	100.275		

Gerlach, 1926

%	b.t.	%	b.t.
3.2	100.2	21.7	101.9
6.3	100.4	25.1	102.4
11.8	100.8		

Tobler, 1855

%		f.t.
19.8		0
27.1		25
32.0		40
35.5	(7+1)	55

Loewel, 1855

f.t.	%		
	(7+1) I	(7+1) II	(6+1)
0	20.63	25.75	28.95
10	23.61	27.91	29.69
20	26.25	30.00	30.50

Mulder, 1866

%	f.t.	%	f.t.
21.20	0	29.08	29.75
22.78	4.25	34.55	55
23.61	9	36.99	68
23.61	9.5	41.59	95
25.20	15.3	42.03	99.7
28.06	25	43.79	108.4

Rüdorff, 1872

%	f.t.	%	f.t.
4.43	-0.7	13.41	-2.60
8.14	-1.35	13.94	-2.80
9.44	-1.65	16.26	-3.75
11.26	-2.15		

de Coppet, 1872

%	f.t.	%	f.t.
4.4	-0.74	18.5	-4.93
8.2	-1.48	20.4	-6.91
11.3	-2.20	21.8	-7.20
14.0	-2.90	23.2	-8.40
16.4	-3.84		

Guthrie, 1876

%	f.t.	%	f.t.	
5	-0.6	21.86	-5.0	E
10	-1.5	21.9	0.0	(7+1)
15	-3.0	25	+15.0	
20	-4.8	30	+31.0	

Etard, 1893-94

%	f.t.	%	f.t.
20.9	2	41.5	94
22.5	7	45.3	130
26.0	23	38.0	145
35.6	67	29.3	164
38.5	81	20.4	188

van 't Hoff, Meyerhoffer and Smith, 1901

E	f.t.
ice + (2+1)	-3.9
(12+1) + (7+1)	+1.8
(7+1) + (6+1)	48.2
(5+1) + (4+1)	77.5

Jones, 1904 and Jones and Getman, 1904

M	f.t.	M	f.t.
0.060	-0.138	0.724	-1.390
.121	-0.263	1.449	-3.240
.242	-0.484	1.932	-5.070
.483	-0.920		

Küpper, 1926

%	f.t.	%	f.t.
20.44(12+1)	0	33.50	50
22.10(7+1)	5	34.50	55
23.60	10	34.90	57
24.98	15.3	35.30	59
25.65	17.9	35.50	60
26.20	20	35.69	61
26.37	20.6	36.25	64
27.50	25	37.38(1+1)	70
28.16	27	37.50	71
29.00	30	38.60	80
29.23	31	40.20	83
30.27	35.5	40.45	90
31.30	40	40.80	100
33.40(6+1)	49		

Benrath and Schröder, 1927

25° = solid phase (7+1)

Smits and Mazee, 1928

mol%	f.t.	mol%	f.t.
	(6+1)		
7.45	60.3	8.72	79.0
7.81	67.4	9.30	85.3
8.19	72.0	9.36	85.7
8.21	72.4	9.65	88.6
	(1+1)		
7.09	95.4	7.50	80
7.16	90	7.16	90
7.50	80	7.09	95.4

Schröder, 1929

mol%	f.t.	mol%	f.t.
0 (8+1)	3.70	6.76	50 (6+1)
10 (7+1)	3.92	7.25	60
20	4.88	8.00	70
30	5.52	8.06	80 (1+1)
40	6.13	7.69	90
		7.52	100

Benrath, 1941

%	f.t.	%	f.t.
5	210	30	210
15	210	35	182
20	210	40	165
25	212	45	125

Kamecki and Palej, 1955

tr.t.	phases
53	(7+1) - (6+1)
87-92	(6+1) - (5+1) and (4+1)
106	(4+1) - (3+1)
122-124	(2+1) - (1+1)
156-160 164-167	(3+2) - (1+1)
320-330	(1+1) - 0 aq.

Chernogorenko, 1956

E_1 : -5.2

E_2 : -3.9

(7+1) (12+1)

Busz, 1938

%	crystallization velocity (m/sec.)
14-16°	
0	1480
5.7	1570
10.8	1650
21.6	1820

Properties of phases.

Anthon, 1835

%	d	%	d
	15°		
0.99	0.0999	23.07	0.1120
1.96	.1006	23.66	.1124
2.91	.1016	24.24	.1128
3.84	.1020	24.81	.1131
4.76	.1024	25.37	.1134
5.66	.1029	25.92	.1137
6.54	.1034	26.47	.1140
7.707	.1039	27.007	.1143
8.25	.1043	27.53	.1145
9.09	.1046	28.05	.1147
9.909	.1050	28.57	.1150
10.71	.1055	29.07	.1153
11.504	.1059	29.57	.1155
12.28	.1064	30.06	.1158
13.04	.1068	30.55	.1161
13.79	.1072	31.03	.1164
14.52	.1075	31.506	.1166
15.25	.1080	31.97	.1168
15.96	.1084	32.43	.1170
16.66	.1088	32.88	.1172
17.35	.1091	33.33	.1174
18.03	.1095	37.50	.1207
18.69	.1098	41.17	.1230
19.35	.1101	44.44	.1250
20.00	.1104	47.36	.1270
20.63	.1107	50.00	.1282 (27.5°)
21.26	.1111	52.38	.1294 (32.5°)
21.87	.1114	54.54	.1309 (37.5°)
22.48	.1117		

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0187	16	1.1691
4	.0393	18	.1929
6	.0599	20	.2175
8	.0808	22	.2429
10	.1021	24	.2694
12	.1238	24.98	.2836
14	.1462		

Bucholz, 1856

%	d	%	d
	18.75°		
2	1.0187	16	1.1709
4	.0394	18	.1941
6	.0604	20	.2177
8	.0818	22	.2417
10	.1035	24	.2660
12	.1256	25	.2783
14	.1481	25.57	.2919

Schiff, 1858 and 1860

%	d	%	d
	23°		
0	0.9976	13.66	1.1398
0.49	1.0023	14.14	.1453
0.98	.0071	14.63	.1508
1.46	.0119	15.12	.1564
1.95	.0168	15.61	.1620
2.44	.0217	16.10	.1675
2.93	.0265	16.58	.1731
3.42	.0314	17.07	.1788
3.90	.0362	17.56	.1846
4.39	.0412	18.05	.1904
4.88	.0461	18.53	.1962
5.37	.0511	19.02	.2020
5.86	.0561	19.51	.2079
6.34	.0611	20.00	.2138
6.83	.0662	20.49	.2198
7.32	.0713	20.98	.2258
7.81	.0764	21.46	.2319
8.30	.0815	21.95	.2380
8.78	.0867	22.44	.2442
9.27	.0918	22.93	.2503
9.76	.0970	23.42	.2565
10.25	.1023	23.90	.2628
10.74	.1076	24.39	.2691
11.22	.1129	24.88	.2755
11.71	.1182	25.37	.2819
12.20	.1234	25.86	.2884
12.69	.1288	26.34	.2948
13.18	.1343		

Gerlach, 1859

%	d	%	d
	15°		
0.49	1.004	13.72	1.145
0.98	.009	14.21	.150
1.47	.015	14.70	.151
1.96	.020	15.19	.162
2.45	.025	15.68	.169
2.94	.030	16.17	.174
3.43	.035	16.66	.180
3.92	.039	17.15	.186
4.51	.050	17.64	.192
4.90	.055	18.13	.198
5.39	.060	18.62	.203
5.88	.065	19.11	.209
6.37	.070	19.60	.215
6.86	.071	20.09	.221
7.35	.081	20.58	.228
7.84	.086	21.07	.234
8.33	.091	21.56	.239
8.82	.096	22.05	.245
9.31	.101	22.54	.252
9.80	.107	23.03	.259
10.29	.113	23.52	.265
10.78	.119	24.01	.271
11.27	.124	24.50	.278
11.76	.129	25.00	.284
12.28	.134	25.49	.290
12.74	.139	25.98	.298
13.23			

Oudemans, 1868

%	d	%	d
	11.2°		
0	0.9996	10.26	1.1071
0.49	1.0046	10.75	.1125
0.98	.0096	11.24	.1179
1.47	.0146	11.73	.1234
1.95	.0196	12.21	.1289
2.44	.0246	12.70	.1344
2.95	.0296	13.19	.1399
3.42	.0346	13.68	.1454
3.91	.0396	14.17	.1510
4.40	.0446	14.65	.1566
4.88	.0497	15.14	.1622
5.37	.0548	15.63	.1679
5.86	.0599	16.12	.1736
6.35	.0650	16.61	.1793
6.84	.0702	17.00	.1850
7.32	.0754	17.58	.1908
7.81	.0807	18.07	.1965
8.30	.0859	18.56	.2023
8.79	.0911	19.05	.2082
9.28	.0964	19.53	.2140
9.77	.1018		

Favre and Valson, 1874

m	d	m	d
	23.5°		
0	0.997	1.5	1.166
0.5	1.059	2.0	1.214
1.0	1.114	2.5	1.260

Kohlrausch, 1879

%	d
15°	
5	1.0510
10	.1052
15	.1602
20	.2200
25	.2861

Rönnberg, 1880

t	d	t	d
	5 %		
3.7	1.04938	15.3	1.04750
10.0	1.04851	20.2	1.04630
	10 %		
3.7	1.09621	15.3	1.09361
10.4	1.09488	20.6	1.09207
	15 %		
3.7	1.14059	15.5	1.13764
10.4	1.13911	20.6	1.13597
	20 %		
3.6	1.18351	15.2	1.18022
10.2	1.18181	20.6	1.17833
	25 %		
3.6	1.22508	15.0	1.22165
10.2	1.22329	20.5	1.21965

Klein, 1886

N	d	N	d
	18°		
0.5	1.0285	2.5	1.1395
1	1.0574	3.423	1.187
1.5	1.0851	4.108	1.222
2	1.1125		

Bindel, 1890

%	d	%	d
	20°		
11.8	1.1253	30.8	1.3024
25.0	.2468	33.9	.3375
27.0	.2657	35.7	.3598
29.4	.2877	37.7	.3845

Charpy, 1893

%	d	%	d
	0°		
0	0.9999	9.6218	1.1071
2.5907	1.0283	11.7458	.1328
5.0447	.0556	13.8000	.1585
7.4046	.0825		

Barnes and Scott, 1898

%	d	%	d
	18.2°		
0	0.9986	18.41	1.1950
2.01	1.0187	21.60	.2330
8.08	.0803	24.53	.2693
11.29	.1147	25.91	.2860
12.63	.1292	26.25	.2903
13.79	.1423		

Stocker, 1920

%	t	d	%	t	d
0	23.8	0.9975	16.92	23.7	1.1762
5.98	21.4	1.0580	17.56	20.3	1.1745
11.42	18.9	1.1158	23.26	22.2	1.2522

Forchheimer, 1900

%	d
20°	
0	0.9982
10.14	0.1049
18.61	1.2019
25.46	1.2879

Brümmer, 1902

%	d	%	d
	15°		
0	0.9993	8.706	1.0914
2.974	1.0306	11.520	1.1216
5.749	1.0597	14.359	1.1550

Jones, 1904; Jones and Getman, 1904

M	d	M	d
	0°		
0.060	1.007040	0.724	1.077856
.121	.014736	1.449	.154336
.242	.027000	1.932	.205724
.483	.054476		

Landesen, 1904

t	d	t	d
0.1248 %			
34.12	0.994440	54.93	0.985848
40.00	.992316	59.50	.983576
45.20	.990237	64.28	.981063
50.11	.988096	71.92	.976796
0.5036 %			
32.54	0.994772	60.14	0.983044
36.56	.993408	64.45	.980773
40.35	.991990	64.48	.980759
45.09	.990082	70.06	.977677
50.14	.987883	75.00	.974802
55.08	.985560	79.97	.971789
1.0047 %			
33.61	0.994202	54.48	0.985662
38.90	.992319	59.16	.983359
44.48	.990120	64.31	.980684
49.84	.987805	71.76	.976590
5.6345 %			
28.11	0.994267	60.67	0.980899
33.58	.992454	60.68	.980894
38.92	.990496	65.41	.978476
44.40	.988323	70.02	.976036
49.84	.985998	74.98	.973295
55.06	.983616	80.02	.970406
11.0 %			
28.62	0.992761	51.35	0.983954
33.37	.991116	57.53	.981173
38.77	.989125	66.68	.976766
44.35	.986921		
12.8455 %			
29.35	0.992178	58.25	0.980580
34.46	.990387	58.30	.980557
39.70	.988420	63.22	.978251
44.48	.986535	68.77	.975529
49.24	.984561	74.35	.972698
54.50	.982273	79.97	.96716

Pann, 1906

%	d 10°.	d 18°	d 30°
0	0.999727	0.998622	0.995673
10.63	1.11297	1.11088	1.10671
18.83	1.21011	1.20595	1.20379
21.35	1.23860	1.23698	-
33.61	-	-	1.29624

Chéneveau, 1907

%	d	%	d
15°			
0	0.9991	9.53	1.0998
3.38	1.0345	12.34	1.1313
6.55	1.0676	14.99	1.1628
8.06	1.0836		
13°			
0	0.9994	6.20	1.0641
2.15	1.0220	8.11	1.0845
4.22	1.0429	9.94	1.1048
5.22	1.0535		

Herz, 1914

N	d
25°	
0	0.9971
1.22	1.0665
2.44	1.1333
3.66	1.1976
4.88	1.2580

Serowy, 1922

%	d
16°	
4.39	1.0027
10.57	1.0093
19.25	1.1669
27.88	1.2795
32.87	1.3532
34.42	1.3733

Manchot, Jahrstorfer and Zepter, 1924

%	d
25°	
10.792	1.0999
10.835	1.0992
21.430	1.1925
21.584	1.1944

Rakshit, 1925

% (7+1)	d
20°	
1	1.00390
5	.02359
10	.04759
30	.13904
50	.22504

Volkmann, 1930

%	d
15-16°	
0	0.999
6.62	1.0680
13.98	.1497
24.25	.2744

Gibson, 1934

%	d
25°	
0	0.9971
5	1.048
10	.102
15	.158
20	.219
25	.289

Okazaki, 1935

%	d
28°	
7.46	1.0732
13.30	.1376
18.40	.1973
23.70	.2632

Guillaume, 1946

%	d
20°	
7.08	1.0720
12.47	.1330
13.3	.1429
25.2	.2877

Urazov and Efimenko, 1956

%	d	%	d
25°			
2.0	1.0173	16.0	1.1698
4.0	.0378	18.0	.1936
6.0	.0587	20.0	.2179
8.0	.0800	22.0	.2428
10.0	.1017	24.0	.2682
12.0	.1239	26.0	.2942
14.0	.1466	26.64	.3025

De Heen, 1881

t	relative volume	t	relative volume
10.56 %			
10.00	1.000000	42.53	1.010260
15.50	.001282	52.81	.014739
20.12	.002528	60.30	.018473
29.26	.005330	70.42	.023886
35.80	.007635		
15.95 %			
10.00	1.000000	42.70	1.010815
15.18	.001575	52.40	.014683
20.80	.003339	60.20	.017936
28.40	.005736	66.12	.020756
35.40	.008125		
21.50 %			
10.00	1.000000	42.80	1.010835
15.40	.001447	50.35	.013976
21.80	.003346	59.20	.018050
28.30	.005443	66.10	.021299
35.80	.008113		

de Lannoy, 1895

t	relative volume			
	1.81 %	4.73 %	9.40 %	14.50 %
0	1.00000	1.00000	1.00000	1.00000
10	.00055	.00121	.00175	.00232
20	.00231	.00312	.00412	.00500
30	.00513	.00605	.00723	.00828
40	.00865	.00968	.01095	.01190
50	.01288	.01382	.01503	.01598
60	.01781	.01867	.01980	.02052
70	.02354	.02410	.02506	.02551
80	.03015	.03040	.03074	.03097

Gibson, 1934

%	π (1-1000 bars)
	25°
0	39.46
5	33.67
10	28.46
15	23.86
20	19.80
25	16.39

Grotrian, 1876

%	t	η	t	η
4.70	21.87	1661	24.99	1524
9.96	21.77	2485	25.86	2210
14.80	21.68	3843	25.97	3393

%	η	$\tau.10^4$
	18°	
1.44	1586	-
3.67	1798	-
7.41	2171	291
11.08	2600	326
14.85	3395	333
19.61	4915	370
22.61	6488	343
29.75	14350	516

Herz, 1914 and 1917

N	η
	25°
0	895
1.220	1339
2.440	2040
3.660	3243
4.880	5309

Urazov and Efimenko, 1956

%	η(water=1)	%	η(water=1)
	25°		
2.02	1.0918	18.04	2.8530
4.01	.1947	20.05	3.3795
6.02	.3210	22.04	4.0586
8.02	.4635	24.03	4.9530
9.99	.6420	24.80	5.3905
10.73	.7103	25.26	5.6754
12.44	.9158	25.79	6.0424
14.02	2.1206	26.19	6.3411
16.03	.4506	26.64	6.6980

Smits and Mazee, 1928

t	flow in sec.	t	flow in sec.
	sat.sol.		
62.5	53.4	82.0	53.2
64.9	54.4	85.0	52.9
68.0	55.3	89.0	52.2
70.0	56.2	94.0	51.3
78.5	54.1		

Stocker, 1920

%	σ	%	σ
	18°		
0	72.51	16.92	76.32
5.98	73.82	17.56	76.33
11.42	74.39	23.26	78.61

Brümmer, 1902

%	σ	%	σ
	15°		
0	74.92	8.706	74.41
2.974	74.63	11.520	80.44
5.749	74.47	14.359	84.49

Forch, 1905

N	t	σ
1.279	15.8	78.00
2.171	16.0	78.85
2.582	15.7	79.40
3.375	15.8	80.41
4.262	15.9	81.91

Pann, 1906

%	σ	
	10°	30°
0	75.37	72.32
10.63	77.45	74.25
18.83	79.74	76.80
21.35	80.68	-
33.60	-	80.54

Volkmann, 1930

%	σ
15-16°	
0	73.2
6.62	75.0
13.98	76.3
24.25	80.2

Semenchenko and Rustamov, 1936

t	σ	t	σ	t	σ
0 M		0.5 M		1 M	
0	75.49	0	76.00	0	76.65
20	72.53	20	73.30	20	74.00
40	69.54	40	70.51	40	71.00
90	66.00	60	67.16	60	67.87
1.5 M		2.0 M			
0	77.16	0	74.80		
20	74.61	20	75.54		
40	71.50	40	77.40		
60	68.19	60	79.20		

Börner, 1869

t	n_{H_α}	t	n_{H_β}	t	n_{H_γ}
		10%			
45.2	1.345980	44.15	1.352320	43.9	1.355748
40.75	.346698	40.25	.352927	39.75	.356283
35.2	.347469	34.75	.353749	33.7	.357317
30.3	.348204	30.45	.354392	30.4	.357709
27.9	.348508	27.8	.354749	27.75	.358136
		30%			
44.5	1.374717	44.2	1.381223	43.7	1.384763
39.5	.375387	40.3	.381890	39.5	.385463
34.35	.376091	34.9	.382626	34.1	.386232
31.25	.376619	31.1	.383152	30.75	.386757
27.8	.377076	27.7	.383677	27.65	.387211

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.060	1.32711	0.724	1.34191
.121	.32829	1.449	.35643
.242	.33110	1.932	.36579
.483	.33684		

Chéneveau, 1907

%	n_D	%	n_D
15°		13°	
0	1.3334	0	1.3335
3.38	.3401	2.15	.3377
6.55	.3463	4.22	.3417
8.06	.3495	5.22	.3437
9.53	.3526	6.20	.3458
12.34	.3585	8.11	.3496
14.94	.3639	9.94	.3536

Guillaume, 1946

%	n_{5780}
20°	
7.08	1.3483
12.47	.3605
13.3	.3615
25.2	.3859

%	$(\alpha)^*$ magn. 10^6 5780 Å
20°	
7.08	3.789
12.47	.646
13.3	.618
25.2	.288

*in radians, gauss, centim.

Studley, 1907

%	χ	%	χ
16-21°			
0	-0.720(20°)	11.53	-0.789
4.73	-0.765	12.40	-0.786
8.95	-0.780	21.04	-0.823
9.48	-0.698 ?		

Forchheimer, 1900

%	rotatory magnetic polarization
20°	
10.14	0.306
18.61	1.313
25.46	0.305

Okazaki, 1935

%	Verdet's constant 3441 Å
	28°
7.46	0.04504
13.30	.04550
18.40	.04584
23.70	.04630

Grotrian, 1876

%	κ	τ.10⁴
	18°	
7.41	223.2	264
11.08	305.9	260
14.85	371.1	263
19.61	421.5	271
22.61	430.7	291
29.75	388.6	360

Berggren, 1877

%	t	κ	%	t	κ
4.41	8.7	184	12.87	7.5	334
6.36	7.8	232	15.00	7.8	351
8.12	8	268	16.25	7.7	355
9.72	7.8	295	18.29	6.3	330
11.16	8	320	19.69	7.5	317

Kohlrausch, 1879

%	κ	τ.10⁴
	18°	
5	263	227
10	412	242
15	478	253
20	474	270
25	413	290

Klein, 1886

N	κ 18°	κ 26°	τ.10⁴
0.5	165.0	195.1	229
1	270.7	320.7	232
1.5	348.5	413.6	234
2	403.5	479.8	237
2.5	437.8	524.3	247
3.423	461.7	-	-
4.108	452.5	-	-

Holland, 1898

t	κ	τ.10⁴
	17.3%	
10	369.2	-
18	457.1	240
20	480.0	260
30	599.8	-
40	723.5	265
50	845.8	266
60	964.9	264
70	1075.0	260
80	1075.4	253

Jones, 1904; Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.060	56.78	0.724	28.06
0.121	52.35	1.449	19.65
0.242	42.36	1.932	14.13
0.483	33.59		

Mc Kie, 1924

N	κ	N	κ
		18°	
0.4356	157.1	1.388	354.1
.5028	175.2	.622	386.5
.7520	236.1	.687	394.3
.9384	275.3	.759	402.2
1.0010	287.7	.837	411.0
.2120	325.2	.922	421.1
.2660	334.7	2.017	430.2
.3240	343.8		

Heim, 1886

t	relative resistance
31.45%	
4	1136.7
5	1087.6
10	877.0
14.5	735.0
20	605.7
25	514.4
30	442.5
35	381.5
40	336.3
45	299.4
50	268.4
55	243.6
60	223.2

Urazov and Efimenko, 1956

%	κ	%	κ
	25°		
2.02	150	17.99	575
4.01	256	18.74	565
6.02	342	23.99	548
8.02	417	24.99	519
9.99	476	24.99	499
12.00	521	25.79	483
14.02	550	26.64	462
16.03	568		

Thermal constants.

Pagliani, 1880

%	U	%	U
0	1.001	16	0.815
11.76	0.862	21.05	0.801
14.28	0.832	25	0.755

Bindel, 1890

%	U	%	U
	20°		
11.8	0.857	30.8	0.697
25.0	.740	33.9	.669
27.0	.730	35.7	.652
29.4	.709	37.7	.633

Teudt, 1900

t	U	t	U
21.66%		10.32%	
32.4	0.7515	31.8	0.8430
33.4	.7615	34.2	.8567
34.5	.7782	35.6	.8580
35.7	.7619	40.0	.8650
41.6	.7768	41.3	.8600
41.4	.7787	46.1	.8712
44.6	.7814	48.9	.8835
47.6	.7870	50.1	.8823
50.8	.7915	53.1	.8899
54.3	.7957	54.3	.8968

Serowy, 1922

mol%	U	mol%	U
16-100°		16-50°	
4.39	0.9350	1.72	0.9763
10.57	.8909	8.81	.9176
19.25	.8181	19.40	.8159
27.88	.7551	22.41	.8038
34.42	.7513		

Gerlach, 1926

%	t	U	%	t	U
3.2	19-52	0.9548	21.7	16.48	0.7510
6.3	18	.9170	25.0	19-24	.7550
11.8	14-18	.8620	25.1	18-90	.7440

Bindel, 1890

%	Q diss.	%	Q diss.
	20°		
11.8	5287	30.8	4942
25.0	5106	33.9	4777
27.0	5041	35.7	4724
29.4	4987	37.7	4617

Kaganovich and Miscenko, 1951

m	Q diss	m	Q diss	m	Q diss
25°		25°		35°	
0.0000	3080	.177	4020	0.055	3852
.0012	3116	.232	4082		
.0042	3260	.313	4164	45°	
.0088	3389	.384	4201		
.0165	3546	.465	4233	0.054	3847
.0383	3762	.603	4325	.068	3860
.054	3838	.770	4370	.110	4051
.061	3850	.864	4428	.121	4105
.091	3883	.948	4470	.139	4120
.113	3996	1.030	4510	.151	4110 ?
.151	4018	1.087	4540	.171	4160

Water + Magnesium acid sulfate ($MgH_2O_8S_2$)

Tammann, 1885

%	p	%	p
	100°		
5.95	749.7	20.69	703.0
10.79	739.1	26.21	673.6
14.95	725.5	32.76	623.9

Water + Magnesium ammonium sulfate ($MgH_8N_2O_8S_2$)

Tobler, 1855

%	f.t.	%	f.t.
5.78(6+1)	0	14.27	45
8.71	10	16.54	50
9.50	15	16.94	55
10.63	20	18.57	60
12.27	30	21.93	75

Porlezza, 1914

%	f.t.	%	f.t.
0	0	10.58	0 (6+1)
1.01	-0.34	12.75	10
1.98	.56	15.23	20
2.98	.785	17.84	30
3.77	.965	20.51	40
4.92	1.23	23.18	50
5.88	.41	26.02	60
6.56	.60	29.07	70
7.42	.83	32.58	80
8.34	2.05	36.03	90
-	2.34 E	39.66	100

Water + Magnesium sodium sulfate ($MgNa_2O_8S_2$)

Roozeboom, 1887 and 1888

f.t.	mol%
30	3.46
35	3.59
47	3.47

Water + Magnesium potassium sulfate ($MgK_2O_8S_2$)

Chernogorenko, 1956

E: -2.8 ice + (6+1)

Tobler, 1855

%	f.t.	%	f.t.
9.04	0	21.09	45
11.99	10	23.39	55
14.63	20	24.45	60
17.06	30	25.34	65
18.24	35	27.38	75
			(6+1)

Schiff, 1860

%	d	%	d
	15°		
0	0.9991	8.78	1.0798
0.73	1.0055	9.51	.0868
1.46	.0120	10.24	.0940
2.19	.0186	10.97	.1011
2.93	.0252	11.97	.1084
3.66	.0318	12.43	.1157
4.39	.0385	13.16	.1230
5.12	.0453	13.89	.1304
5.85	.0520	14.63	.1378
6.58	.0580	15.36	.1453
7.32	.0658	16.09	.1529
8.05	.0727		

Buchholz, 1856

%	d	%	d
	18.75°		
2	1.0162	16	1.1500
4	.0343	18	.1704
6	.0527	20	.1913
8	.0715	22	.2124
10	.0906	24	.2338
12	.1101	25	.2448
14	.1298	26.34	.2609

Schmelzle and Westfall, 1944

N	d	η (water=1)
	25°	
0.0422	0.9995	1.0139
.0866	1.0034	.0210
.1696	.0099	.0487
.2506	.0158	.0643
.3992	.0258	.1059
.5366	.0372	.1551
.6582	.0479	.2003
.8042	.0571	.2393
1.0238	.0728	.3160
1.5082	.1072	.4940
2.7160	.1917	2.1960
4.4960	.3106	4.4660

Water + Magnesium sulfamate ($MgH_4N_2O_6S_2$)

King and Hooper, 1941

%	f.t.	%	f.t.
50.21	0	57.60	50
50.84	10	59.54	60
51.64	20	62.00	70
53.34	30	63.86	75
54.76	40	66.36	80

Water + Magnesium thiosulfate (MgS_2O_3)

Silberman and Ivanov, 1946

%	f.t.	%	f.t.
30.69	0 (6+1)	38.91	61.25
31.97	10.5	41.18	73.25
34.51	28	41.45	75.75
35.48	36.5	42.80	80.75
36.10	43.5	44.81	85.25
36.64	46.5	45.0	86.75
37.62	54		

Water + Magnesium benzenesulfonate ($MgC_{12}H_{10}O_6S_2$)

Ephraim and Seger, 1925

c	f.t.	c	f.t.
7.496	17	20.420	65
11.153	36	26.045	80.5
14.569	50	26.804	82

Water + Magnesium formate ($MgC_2H_2O_4$)

Ashton, Houston and Saylor, 1933

%	f.t.	%	f.t.
12.28	0	15.25	60
12.34	10	16.00	70
12.56	20	17.00	80
13.04	30	18.15	90
13.74	40	19.35	100
14.46	50		

Water + Magnesium acetate ($MgC_4H_6O_4$)

Goode, Bayliss and Rivett, 1928

m	p	m	p
	25°		
0.4185	22.90	2.703	19.33
0.800	22.73	3.430	19.05
1.026	22.40	3.645	19.11
1.205	21.86	4.331	19.83
1.650	21.03	4.55	20.36
1.784	20.71		

Stokes, 1946

isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1431	0.1847	1.053	1.443
0.2013	0.2597	1.259	1.740
0.2503	0.3245	1.728	2.439
0.3799	0.4930	2.205	3.173
0.4286	0.5588	2.748	3.994
0.8053	1.337	3.300	4.794
0.9605	1.355	4.149	5.984

m_1 = m of $MgC_4H_6O_4$ m_2 = m of NaCl

Rubien, 1911

N	d	n_D	N	d	n_D
		18°			
0	0.99862	1.33327	1.191	1.04538	1.34840
0.0969	1.00263	.33456	2.391	.0899	.36287
.243	.00844	.33649	3.713	.1374	.37800
.603	.02256	.34108			

Heydweiller, 1912

N	d	κ	N	d	κ
		18°			
0.0969	1.00263	50.2	1.191	1.04538	266.4
.2427	.00844	103.6	2.391	.0899	284.3
.603	.02256	192.0	3.713	.1374	213.0

Guillaume, 1946

%	d	$(\alpha)^*$ magn. 10^6	n 5780Å
		20°	
9.24	1.0540	3.874	1.3516
13.70	.0812	.825	.3599
20.03	.1150	.749	.3708
21.90	.1262	.729	.3742

* in radians, gauss, centim.

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34604
0.1	.33461	2.0	.35826
0.2	.33593	3.0	.36995
0.5	.33977		

Water + Stannous chloride ($SnCl_2$)

Gerlach, 2867

%	d	%	d
	15°		
0.84	1.006	32.75	1.318
1.68	.012	33.59	.3287
2.52	.019	34.43	.340
3.36	.025	35.27	.351
4.20	.0322	36.11	.362
5.04	.039	36.95	.373
5.88	.046	37.79	.3838
6.72	.053	38.63	.396
7.56	.060	39.47	.408
8.40	.0674	40.30	.420
9.24	.075	40.14	.432
10.08	.082	41.98	.4439
10.92	.089	42.82	.457
11.76	.096	43.66	.470
12.60	.104	44.50	.483
13.44	.112	45.34	.496
14.28	.120	46.18	.5093
15.12	.127	47.02	.524
15.96	.135	47.86	.538
16.79	.1432	48.70	.553
17.63	.151	49.54	.567
18.47	.160	50.38	.5807
19.31	.168	51.22	.597
20.15	.176	52.06	.613
20.99	.1845	52.90	.628
21.83	.193	53.74	.643
22.67	.202	54.58	.6584
23.51	.211	55.42	.676
24.35	.220	56.26	.693
25.19	.229	57.10	.710
26.03	.239	57.94	.726
26.87	.248	58.78	.7437
27.71	.258	59.62	.762
28.55	.267	60.46	.781
29.39	.2768	61.29	.800
30.23	.287	62.14	.819
31.07	.298	62.98	.8383
31.91	.308		

Chéneveau, 1907

%	d	n_D
	15°	
0	0.9991	1.3335
18.55	1.1114	1.3591
33.73	1.2217	1.3855

Water + Stannous nitrate (SnN_2O_6)

Guthrie, 1876

%	f.t.	%	f.t.
5	-0.5	20	-3.8
10	-1.2	26	-6 E
15	-2.3	29.62	0.0

Water + Lead chloride ($PbCl_2$)

Morey and Chen, 1956

f.t.	P	f.t.	P
	V + L + C		
374	207	400	164

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
4.74	141	14.3	257
5.37	150	15.4	266
5.67	160	16.7	276
7.39	187	18.4	287
7.93	191	24.76	345
8.33	193	81.2	351
8.47	195	87.8	362
11.3	230	94.4	407
12.3	242	100.0	499

Water + Lead bromide ($PbBr_2$)

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
7.3	143	33.5	285
12.0	181	46.7	302
17.1	217	57.5	302
22.4	245	71.0	302
23.7	250	86.6	307
28.2	267	100.0	373

%	crit.t.	%	crit.t.
0	370	28.6	355
10.0	369	37.9	330
18.2	362		

Water + Lead iodide (PbI_2)

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
1.42	175	12.0-82	334
3.65	234	87.5	338
6.37	280	95.3	360
6.44	281	100.0	412
8.43	291		

Water + Lead potassium iodide (K_2PbI_4)

Schreinemakers, 1892

mol %	f.t.	mol %	f.t.
12.36	157	20.04	194
14.93	172	23.43	201
17.70	186		

Water + Lead nitrate (PbN_2O_6)

Heterogeneous equilibria .

Tammann, 1885

%	p	%	p
	100°		
16.65	745.2	44.32	707.7
28.82	730.8	52.86	690.6
41.37	712.2		

Gerlach, 1885

%	b.t.	%	b.t.
0	100	39.32	102
9.91	100.5	46.52	102.5
20.63	101	52.60	103
30.55	101.5	57.81	103.5

Buchanan, 1899

%	b.t.	%	b.t.
749.2mm		741.2mm	
34.31	101.02	14.64	99.81
38.43	101.22	17.22	99.91
41.62	101.42	20.01	100.01
44.33	101.61	22.51	100.11
46.83	101.84	25.10	100.21
48.98	102.04	27.25	100.31
51.30	102.24	31.32	100.51
53.97	102.54	35.11	100.72
56.16	102.75	38.06	100.92
56.98	102.89	40.88	101.12
		44.33	101.32
		47.78	101.63

Kopp, 1840

%	f.t.
36.75	22.6
37.15	22.0
38.69	24.7

Kremers, 1854

%	f.t.	%	f.t.
27.94	0	50.25	65
32.57	10	53.54	85
37.73	25	58.14	100
44.44	45		

Mulder, 1864

%	f.t.	%	f.t.
26.74	0	44.81	51
28.11	4.25	51.20	78
29.48	7	52.40	81
30.02	9	54.30	91
33.15	17	56.79	104.7 b.t.
34.98	21		

Rüdorff, 1872

%	f.t.	%	f.t.
7.41	-0.8	21.88	-2.20
9.09	-1.0	23.08	-2.30
13.79	-1.45	24.24	-2.40
16.67	-1.70	25.37	-2.55
19.35	-1.95		

de Coppet, 1872

%	f.t.
7.41	- 0.9
10.71	- 1.2
13.79	- 1.5
19.35	- 2.0
24.24	- 2.5

Guthrie, 1876

%	f.t.	%	f.t.
5	-0.3	20	-1.2
10	-0.5	26.23	-2.5 E
15	-0.7	29.89	0.0

Dupuy, 1884

%	f.t.
30.12	15
42.88	30
54.54	60
56.48	100

Properties of phases . Density .

Kremers, 1855 - 1856

%	d	%	d
	19.5°		
0	0.998	20	1.201
5	1.043	25	1.264
10	1.091	30	1.332
15	1.142	35	1.412

Fouqué, 1867

%	d°	t	d
1.08	1.0098	13	1.0096
4.17	1.0401	10	1.0399
12.20	1.1358	13	1.1339

Forster, 1878

%	t	d
0	24.0	0.9973
14.27	25.2	1.1400
31.65	24.0	1.3695

Long, 1880

%	d	%	d
	15°		
5	1.0449	20	1.2043
10	.0937	25	.2678
15	.1467	30	.3358

Gerlach, 1887

%	d	%	d
	17.5°		
0	0.998	20	1.198
5	1.043	25	.261
10	.091	30	.331
15	.143	35	.407

Bindel, 1890

%	d	%	d
	20°		
8.4	1.0771	42.3	1.6045
37.9	1.4854	47.8	1.6663

Traube, 1895

%	d
15°	
0	0.99913
8.030	1.07168
9.349	1.08460
16.042	1.15478

Le Blanc and Rohland, 1896

%	d
20°	
0	0.9982
15.93	1.1537
30.57	1.3413
30.69	1.3442

Schiff, 1896 and 1897

%	d	%	d
	17.5°		
0	0.9987	20	1.1773
1	1.0067	21	.1887
2	.0150	22	.2000
3	.0234	23	.2116
4	.0318	24	.2235
5	.0403	25	.2356
6	.0489	26	.2479
7	.0577	27	.2604
8	.0668	28	.2731
9	.0761	29	.2859
10	.0855	30	.2990
11	.0949	31	.3123
12	.1045	32	.3259
13	.1143	33	.3399
14	.1243	34	.3541
15	.1344	35	.3684
16	.1448	36	.3830
18	.1554	37	.3978
19	.1662	38	.4128

Cheneveau, 1907

%	d	%	d
	13.5°		
0	0.99935	16.19	1.1609
5.94	1.0547	20.65	.2119
11.31	.1076	24.73	.2644
13.81	.1345		

%	t	d
30.10	16.3	1.3363

Herz, 1914

N	d
25°	
0	0.9971
0.682	1.0919
1.363	.1851
2.044	.2786
2.726	.3708

Herz, 1917

N	d
25°	
0	0.9971
0.682	1.0924
1.363	.1846
2.044	.2785
2.726	.3707

Heydweiller, 1921

N	d
18°	
0.2	1.0287
0.5	.0709
1	.1408
2	.2790

Rakshit, 1925

c	d
20°	
1	1.0052
5	.03976
10	.08154
30	.24739
50	.30825

Okazaki, 1935

%	d
28°	
7.69	1.0660
12.16	.1106
17.44	.1685
21.87	.2218
26.48	.2795

Guillaume, 1946

%	d
20°	
11.04	1.1064
17.73	.1756
22.20	.2326
33.19	.3877

de Lannoy, 1895

t	relative volume			
	4%	10%	15%	25%
0	1.00000	1.00000	1.00000	1.00000
10	.00065	.00142	.00170	.00264
20	.00246	.00364	.00427	.00581
30	.00533	.00884	.00765	.00967
40	.00902	.01065	.01174	.01410
50	.01350	.01517	.01650	.01890
60	.01836	.02026	.02172	.02445
70	.02392	.02600	.02750	.03021
80	.03012	.03230	.03380	.03660

Viscosity and surface tension

Herz, 1914 and 1917

N	η(water=1)
25°	
0	1
0.682	1.070
1.363	.172
2.044	.294
2.726	.439

Wagner, 1883

%	η(water°=100)			
	15°	25°	35°	45°
17.93	74.04	59.13	48.52	40.33
32.22	91.85	72.49	59.60	50.56

Malquori, 1929

t	η	t	η
	0.4508 N		
13.6	1265	45.4	675
34.4	807	60.0	680
	0.9017 N		
14.0	1338	45.6	708
35.8	819	60.0	711
	1.3527 N		
12.8	1415	45.6	739
33.8	893	60.0	740
	1.7814 N		
13.5	1551.8	60.0	772.1
	2.2988 N		
13.4	1729.8	60.0	829.6
	2.634 N		
12.9	1806	60.0	847.8

Forch, 1905

M	σ
16°	
0.515	77.01
0.780	77.29
0.915	77.44
1.024	77.46
1.307	77.76

Optical and electrical properties.

Forster, 1878

%	t	n_D
0	24.0	1.3326
14.27	25.2	.3511
31.65	24.0	.3808

Walter, 1889

%	n_D	%	n_D
		15°	
0	1.3334	18.5	1.3567
4.0	1.3381	32.3	1.3792
8.6	1.3435		

Chèneveau, 1907

%	n_D	%	n_D
	13.5°		
0	1.33365	16.19	1.3546
5.94	.3407	20.65	.3616
11.31	.3477	24.73	.3688
13.81	.3512		

t	%	n C	n Tl	n F	n G'
16.3	30.10	1.37529	1.37966	1.38339	1.38788

Barbier and Roux, 1890

c	t	dispersive power
10.0	12.5	0.378
20.0	12.6	.411
30.0	10.1	.440
40.0	12.5	.471

Le Blanc and Rohland, 1896

%	n_D	%	n_D
		20°	
0	1.3333	30.57	1.3779
15.93	1.3536	30.69	1.3789

Guillaume, 1946

%	n 5780Å	(α)*magn.10^6
	20°	
11.04	1.3483	3.821
17.73	.3581	.717
22.20	.3653	.652
33.19	.3859	.486

*= in radians, gauss, centim.

Okazaki, 1935

%	Verdet's constant.10^5
28°	
7.69	4401 (3514 Å)
12.16	4553
17.44	4744
21.87	4912
26.48	5099

Malquori, 1929

t	κ	t	κ	t	κ
0.4508N		0.9017N		1.3527N	
13.6	228.0	14.0	363.4	12.8	461.0
34.4	372.3	35.8	590.7	33.8	766.7
45.4	442.3	45.6	694.0	45.6	898.7
60.0	552.6	60.0	875.4	60.0	1126.0

Heydweiller, 1921

N	λ
18°	
0.2	67.7
0.5	53.4
1	42.3
2	30.7

Long, 1880

%	κ	τ.10^4
	18°	
5	182	238
10	306	251
15	407	251
20	495	250
25	570	252
30	635	257

Bindel, 1890

%	U	Q diss
	20°	
8.4	0.919	-7104
37.9	.646	-4504
42.3	.6045	-4246
47.8	.6663	-4086

Jauch, 1921

N	U
18°	
0.2	0.9621
0.5	.9179
1	.8559
1.5	.8045
2	.7566

Water + Lead perchlorate ($PbCl_2O_8$)

Willard and Kassner, 1926

81.472 %   f.t. = 25° (3+1) d^{25} = 2.7753

Water + Lead acetate ($PbC_4H_6O_4$)

Tammann, 1885

%	p	%	p
100°			
22.23	747.0	50.58	725.4
30.44	742.6	61.03	711.5
39.55	736.4	65.93	701.1

Gerlach, 1886

%	b.t.	%	b.t.
0	100	90.25	109
25.93	100.5	91.32	110
44.13	101	92.42	111
55.56	101.5	93.13	112
63.10	102	93.89	113
68.50	102.5	94.52	114
72.60	103	95.01	115
75.90	103.5	95.51	116
78.49	104	95.92	117
80.58	104.5	96.34	118
82.30	105	96.71	119
83.61	105.5	96.96	120
84.82	106	98.37	125
85.75	106.4	99.45	130
86.96	107	satd	133
88.81	108		

Gerlach, 1867

%	d	%	d
15°			
0	0.9991	22.30	1.1859
0.858	1.0055	23.15	.1945
1.715	.0118	24.01	.2029
2.573	.0182	24.87	.2115
3.43	.0246	25.73	.2200
4.29	.0310	26.58	.2292
5.15	.0377	27.44	.2384
6.00	.0444	28.30	.2475
6.86	.0511	29.16	.2567
7.72	.0578	30.01	.2658
8.58	.0645	30.87	.2757
9.43	.0716	31.73	.2856
10.29	.0787	32.59	.2955
11.15	.0858	33.44	.3053
12.01	.0929	34.30	.3152
12.86	.1000	35.16	.3257
13.72	.1074	36.02	.3367
14.58	.1149	36.88	.3470
15.44	.1224	37.73	.3576
16.29	.1299	38.59	.3683
17.15	.1374	39.65	.3798
18.01	.1454	40.30	.3913
18.87	.1534	41.16	.4029
19.72	.1614	42.02	.4144
20.58	.1694	42.88	.4259
21.44	.1774		

Oudemans, 1868

%	d	%	d
14°			
0	0.9993	14.58	1.1167
0.858	1.0057	15.44	.1243
1.715	.0121	16.29	.1321
2.573	.0186	17.15	.1399
3.43	.0251	18.01	.1478
4.29	.0317	18.87	.1559
5.15	.0384	19.72	.1641
6.00	.0452	20.58	.1724
6.86	.0520	21.44	.1808
7.72	.0589	22.30	.1894
8.58	.0659	23.15	.1981
9.43	.0729	24.01	.2069
10.29	.0800	24.87	.2158
11.15	.0872	25.73	.2248
12.01	.0945	26.58	.2339
12.86	.1018	27.44	.2432
13.72	.1092	28.30	.2525

Rubien, 1911 and Heydweiller

N	d	N	d
18°			
0	0.99862	0.8030	1.09357
0.0811	1.008391	1.598	1.18522
0.1615	1.01798	2.292	1.26508
0.4024	1.04631	3.425	1.39713

Rakshit, 1925

c	d
20°	
1	1.00352
5	.02828
10	.05863
30	.18090

Guillaume, 1946

%	d
20°	
17.76	1.1471
28.55	1.2571

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.8030	1.34736
0.0811	.33477	1.598	.36090
.1615	.33622	2.292	.37258
.4024	.34042		

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	0.5	1.34213
0.1	.33511	1.0	.35075
0.2	.33691	2.0	.36770

Heydweiller and Grube, 1916

w.l. Å	Dn(sol.-aq.).10^5 0.4971N	0.9725N	2.006N
		18°	
2981	1140	2133	4376
3256	1033	1997	4086
3405	1007	1946	3980
3612	983	1839	3864
4679	913	1747	3573

Guillaume, 1946

%	$(\alpha)^*$ magn.10^6	n 5780Å
	20°	
17.76	3.900	1.3562
28.55	3.862	1.3719

*= in radians, gauss, centim.

Blythswood and Marchant, 1899

M	a*
1	2.0
0.5	3.4
.165	5.4
.125	7.4
.100	9.3

* a= thikness in mm of a layer eouivalent to 20 mm of water (absorption of X rays)

Heydweiller, 1912

N	ϰ	N	ϰ
	18°		
0.0811	18.76	1.598	75.1
.1615	28.35	2.292	77.6
.4024	45.98	3.425	73.5
.8030	61.44		

Marignac, 1876

%	U
18-51°	
8.27	0.9327
15.28	.8808
26.51	.7939
41.91	.6824

Water + Lead fluosilicate ($PbSiF_6$)

Jatlov and Pinaevskaya, 1938

%	f.t.	%	f.t.
0	65.48	60 (2+1)	80.11
20	68.97	65	80.75
50	74.16	80	81.06
55	75.38	82.25	100
57	77.70		

LI. WATER + SALTS OF ALCALINE EARTH GROUP.

Water + Calcium chloride ($CaCl_2$)

Heterogeneous equilibria.

Wüllner, 1860

t	p 0%	13.64%	25.50%	45.50%
16.20	13.710	13.212	12.738	10.915
18.40	15.747	15.050	14.651	12.454
19.84	17.212	16.515	15.867	13.579
20.40	17.826	17.730	16.384	14.193
21.85	19.479	18.383	17.638	15.298
23.28	21.272	19.977	19.132	16.594
26.05	25.058	23.866	22.830	19.834
28.35	28.684	27.296	26.105	22.632
32.22	35.800	34.215	32.725	28.358
34.90	41.595	39.960	38.125	33.360
37.50	48.261	46.079	43.893	38.430
38.93	51.866	49.385	47.400	41.140
40.66	56.708	54.030	51.752	45.150
43.05	64.496	61.519	58.359	51.496
45.55	73.274	70.298	67.338	58.752
48.69	86.210	83.092	79.279	69.360
51.70	100.078	96.116	91.959	80.035
53.00	106.636	102.476	98.324	85.529
55.69	121.514	117.155	112.267	97.932
58.30	137.458	132.509	126.986	110.694
58.70	140.062	134.974	128.513	111.680
60.00	148.791	143.514	137.541	120.093
60.60	153.019	147.643	104.974	123.662
62.28	165.456	159.191	151.813	132.200
63.13	171.583	165.417	157.981	138.330
64.10	179.537	176.124	165.741	144.640
65.64	191.075	183.482	175.100	151.939
68.45	217.902	210.113	200.954	175.180
70.60	239.273	230.852	220.571	192.571
72.20	256.287	247.655	236.415	206.976
74.38	281.908	272.657	260.738	228.752
76.83	311.429	300.483	287.417	251.954
78.80	337.747	326.062	312.235	273.511
80.60	363.427	351.489	336.419	295.245
82.63	393.431	380.222	363.969	318.933
84.80	429.516	415.097	396.421	348.160
86.50	459.212	443.361	425.990	372.988
87.65	480.175	464.175	443.954	389.703
89.93	524.150	506.098	484.150	427.868
92.20	571.031	552.288	528.193	468.730
94.65	625.630	605.282	578.777	511.913
96.85	678.310	655.962	627.454	554.587
99.30	741.280	718.280	687.332	607.766

Tammann, 1885

%	p	%	p
	100°		
4.80	744.7	31.34	505.7
10.97	715.8	31.79	499.3
15.93	683.6	34.89	452.2
20.31	645.1	37.76	409.8
25.99	581.5	40.69	366.0
28.53	548.3	41.91	349.3

t	p 0%	7.48%	12.51%	14.34%	23.11%
45.5	36.38	43.0	40.9	40.5	35.6
54.2	39.61	51.0	48.6	48.1	42.4
88.1	49.01	84.3	80.8	79.1	68.9
104.7	51.51	100.2	95.9	94.3	82.2
127.4	56.59	122.3	136.9	115.3	100.5
136.6	58.07	131.3	126.0	124.0	108.0
156.5	61.00	150.3	144.6	142.3	124.7
184.4	64.61	178.3	170.2	167.1	146.8
221.9	68.80	214.0	205.6	202.1	177.1
243.2	70.91	234.8	224.9	221.0	214.4
273.2	73.64	264.1	253.4	248.8	238.3
306.0	76.34	296.1	284.6	279.0	244.7
347.8	79.46	336.0	322.9	317.1	279.0
375.7	81.37	363.5	349.4	342.8	301.3
408.8	83.49	395.2	380.1	372.3	328.8
453.2	86.12	438.6	421.7	414.0	364.9
496.4	88.48	479.9	461.6	453.2	399.8
550.8	91.22	532.6	512.1	503.2	444.0
604.7	93.72	584.4	561.0	551.7	486.8
704.7	97.90	679.5	654.5	644.1	569.2
765.1	100.19	749.1	711.4	700.3	620.0

Emden, 1887

t	p	t	p
	10.15 %		
19.64	16.0	61.38	146.8
24.19	21.0	66.08	182.3
39.79	31.0	69.89	214.55
34.74	38.5	75.44	273.25
40.85	53.5	80.54	337.6
45.19	67.2	85.23	407.3
50.04	85.9	90.30	495.8
54.36	105.8	96.94	633.35
	13.12 %		
20.26	16.15	61.19	143.50
29.55	21.25	65.38	173.45
31.19	30.90	72.00	232.20
34.81	37.80	75.86	274.00
39.44	49.05	80.57	332.30
44.14	62.55	85.69	409.60
51.56	90.85	90.42	491.00
56.80	117.20	93.50	551.60
	19.752 %		
19.62	15.00	60.02	131.05
24.27	19.95	64.79	162.40
29.48	26.95	71.52	219.10
35.05	36.80	76.30	267.80
39.85	47.85	79.78	309.45
44.35	60.65	85.44	387.90
50.56	83.10	90.30	469.15
55.62	106.45	95.21	563.70
	19.990 %		
17.54	13.45	59.55	127.30
23.60	19.35	64.38	159.10
29.28	26.90	72.12	224.95
35.24	37.55	74.22	245.80
39.45	47.10	80.95	324.20
44.20	60.35	86.49	405.00
50.17	81.60	90.08	466.20
54.98	103.85	93.62	533.00
	16.31 %		
20.71	16.00	59.68	128.50
24.98	20.75	65.50	167.90
29.98	27.70	69.97	203.20
35.33	37.40	75.64	259.60
39.42	46.60	80.56	318.60
45.18	63.10	85.35	386.30
50.18	81.30	90.64	473.00
59.26	99.50	95.48	567.00

Bremer, 1888

t	Dp		
	7.480%	14.88%	19.52%
20.90	-0.486	-1.382	-2.168
26.66	0.810	2.089	3.234
31.87	1.076	3.048	4.697
37.15	1.479	4.102	6.388
42.34	2.155	5.660	8.705

t	Dp	
	12.81%	19.61%
23.38	-1.363	-2.388
29.50	2.116	3.923
34.60	2.819	5.494
39.94	3.847	7.508
45.69	5.264	10.438
52.25	7.215	14.624
57.38	8.795	18.381

t	Dp		
	6.578%	10.57%	14.91%
14.93	-0.2005	-0.550	-0.750
20.49	.441	0.914	1.271
25.98	.643	1.322	1.916
30.30	.859	1.704	2.544
34.48	1.216	2.079	3.501
35.82	.337	2.471	3.681
38.82	.597	2.950	4.415
41.39	.793	3.280	5.119
46.31	2.217	4.194	6.670
51.23	2.807	5.407	8.324
56.31	3.941	7.604	11.769
57.10	4.144	7.658	11.931
62.81	5.257	9.830	15.321
67.81	5.732	11.342	18.549

Dieterici, 1898

M	p
0°	
0	4.579
0.982	4.415
1.996	4.137
4.976	3.296
10.44	2.277

Geller, 1911

t	p	t	p
2.91%		4.53%	
18.00	15.083	19.03	15.566
20.30	17.064	21.83	18.296
22.80	19.892	24.43	21.066
25.17	23.042	27.22	24.695
27.30	25.537	29.60	28.925
29.89	30.136	31.90	32.172
32.78	35.584	34.45	37.638
35.21	41.116	37.20	43.950
37.40	45.947	39.30	50.206
39.57	52.100	41.30	57.115
41.47	56.883	44.40	64.906
43.43	63.470	46.60	72.609
45.22	70.639	48.40	81.090
47.10	78.352	50.49	88.608
49.18	85.332		
51.70	96.688		
12.74%		18.26%	
19.97	15.503	21.67	16.004
22.20	17.557	23.60	17.675
24.67	20.674	26.10	20.530
26.90	23.755	28.20	23.261
28.90	29.614	30.56	26.456
33.2	33.089	33.02	29.532
35.5	38.458	35.0	34.249
37.7	43.719	37.2	38.552
39.85	48.919	39.0	42.839
42.0	55.099	40.9	47.487
44.2	60.779	43.3	53.727
45.9	67.165	45.1	59.631
48.2	75.230	46.9	64.722
49.75	81.437	48.3	69.748

Perman and Price, 1912

c	p	c	p
70°		90°	
5.01	231.2	2.75	518.98
7.62	228.0	4.95	514.5
9.97	224.9	7.37	510.8
12.47	221.5	10.64	502.0
17.22	212.7	13.69	490.8
21.51	203.7	18.92	468.5
35.98	165.5	27.97	421.4
50.76	117.8	39.06	351.4
57.76	99.6	57.69	237.0
58.96	94.2	74.12	162.1
80.15	66.0		

Pieper, 1917

t	p	t	p
5.31%		20.226%	
18.0	14.78	18.0	13.34
25.0	22.33	25.0	19.98
29.0	27.90	29.0	25.25
36.0	41.62	36.0	37.76
41.0	54.49	41.0	49.45
47.0	74.51	47.0	68.01
48.0	78.06	48.0	71.29

Baker and Waite, 1921

t	p	t	p
	20.59 %		
58.02	113.3	85.45	370.8
65.90	162.7	90.58	452.7
70.90	203.0	95.24	539.5
76.19	254.2	100.00	641.8
81.00	310.0	103.8	738.6
	40.77 %		
72.66	115.1	100.37	370.5
80.05	161.0	105.56	451.4
85.20	200.8	110.30	537.7
90.70	251.9	114.96	634.2
95.68	308.8	119.35	738.4
	50.25 %		
82.70	112.9	110.61	371.3
90.28	162.4	115.90	451.7
95.45	202.6	120.65	537.5
100.85	253.0	125.39	635.4
106.10	311.6	129.68	738.8
	58.51 %		
90.95	108.7	120.61	362.3
99.00	153.6	126.12	442.4
104.51	193.9	131.20	528.9
110.15	243.7	136.37	632.0
115.42	298.7	140.32	733.5
	sat.sol.		
100.05	118.8	132.32	363.3
105.25	154.6	139.50	440.1
111.90	194.6	144.55	498.7
115.20	219.2	147.90	536.2
118.76	242.9	152.75	596.8
125.58	299.8	173.40	742.5

Ebert, 1930

%	p 0°	p 5.2°	p 18°	p 20°
0	4.58	6.64	15.48	17.54
10	4.30	6.25	14.40	16.55
20	3.75	5.45	12.50	14.40
30	2.85	4.20	9.70	11.30
40	-	-	6.15	7.40
satd	1.80	2.50	5.45	6.15

Hepburn, 1932

m	%	p	m	%	p
		25°			
0.199	2.16	23.50	1.504	14.30	21.61
.284	3.05	23.43	1.985	18.06	20.47
.344	3.65	23.43?	2.495	21.69	19.26
.494	5.20	23.28	2.990	24.91	17.86
.635	6.59	22.12	4.004	30.77	14.73
.796	8.12	22.88	5.826	37.25	9.48
1.000	9.99	22.61	7.278	44.68	7.00
				(satd)	

Lannung, 1936

m	p	m	p
	18°		
3.267	11.09	5.389	6.97
3.441	10.76	6.537	5.22
3.676	10.28	6.855	4.81
4.280	9.07	6.939	4.76
4.409	8.08	7.918	3.81
4.696	8.32	8.985	3.14
4.839	7.97	9.239	2.98

Harrison and Perman, 1927

%	p	%	p
	40°		
0	55.2	19.10	48.2
2.26	55.1	20.52	46.5
2.92	55.0	23.15	45.5
4.01	54.9	27.66	40.2
6.69	54.6	32.79	34.7
10.01	52.2	38.50	27.8
12.57	52.0	44.70	20.7
14.66	50.4	50.90	13.9
16.27	49.7	56.30	11.1
	50°		
0	92.4	22.66	75.3
3.10	90.4	28.32	66.3
5.50	90.2	29.01	65.0
7.58	88.9	33.90	55.4
9.01	88.5	36.60	48.6
11.59	86.8	41.50	38.4
12.72	86.2	46.40	30.5
13.94	85.7	51.40	22.7
16.24	83.0	56.30	16.3
19.00	80.4		
	68°		
0	149.3	22.09	124.0
1.94	149.3	23.70	119.3
2.25	149.2	26.24	112.8
3.18	143.7	28.10	108.9
5.51	146.8	30.40	100.3
6.48	145.7	32.40	94.8
6.69	145.0	36.45	81.5
7.38	144.0	43.50	59.6
10.36	141.3	47.66	48.0
12.38	139.5	51.86	38.4
19.93	129.0	56.1	28.7
	70°		
0	233.8	27.22	175.6
3.39	233.0	30.36	160.5
5.75	228.0	35.54	134.0
7.61	226.3	37.18	125.0
8.08	227.0	40.86	107.2
10.11	222.8	44.10	94.0
12.92	217.2 (sic)	48.36	76.5
15.85	211.4	52.19	60.0
19.81	201.7	56.82	43.2
24.13	186.8	57.92	42.1
25.74	181.2		
	80°		
0	355.4	22.55	295.2
2.48	352.1	24.07	287.6
2.93	351.8	27.82	264.4
3.61	351.0	31.30	240.9
5.73	346.2	37.08	198.9
5.87	346.0	40.81	170.6
8.15	342.9	44.48	143.4
11.60	334.6	49.70	112.8
14.13	318.8	52.96	96.4
15.40	313.7	57.32	75.5
18.17	316.0	60.52	63.2
20.20	304.1		

Roozeboom, 1889

t	p	mol%	
-55	0	6.45	ice + (6+1) + L+V
+29.2	5.67	15.60	(6+1) + (4+1)II + L+V
29.8	6.80	14.10	(6+1) + (4+1)I + L+V
38.4	7.88	17.20	(4+1)II + (2+1) + L+V
45.3	11.77	17.50	(4+1)I + (2+1) + L+V
175.5	842.00	32.60	(2+1) + (1+1) + L+V
above	1 atm.	35.70	(1+1) + 0 aq. + L+V

%	t	p	%	t	p
		C + V + L			
0	0	4.63	25.09	30	-
9.5	--5	3.06	27.80	40	-
14.53	10	2.03	29.83	55	0.0
		(6+1) + L + V			
29.83	55	0.0	45.05	25	6.696
33.33	25	-	47.50	28.5	7.020
35.48	-10	0.972	48.85	29.5	6.910
37.30	0	1.944	50.65	30.2	6.696
39.32	+10	3.456	52.13	29.6	5.830
42.70	20	5.616	53.00	29.2	5.670
		(2+1) + L + V			
56.13	40	8.5	61.39	100	145.0
56.56	45.3	11.77	62.39	110	204.0
56.93	50	15.5	63.77	125	326.0
57.36	55	20.5	64.91	135	435.0
57.79	60	26.5	65.64	140	497.0
58.17	65	34.0	68.20	155	680.0
58.62	70	43.0	68.99	160	744.0
59.06	75	54.0	70.24	165	790.0
59.50	80	66.5	71.83	170	834.0
59.97	85	82.5	74.80	175.5	842.0
60.39	90	100.6			
		(4+1)I + L + V			
47.64	20	4.744	51.75	35	8.640
48.98	25	5.724	53.55	40	10.370
50.14	29.8	6.800	56.56	45.3	11.770
		(4+1)II + L + V			
51.10	20	3.564	53.27	30	5.830
52.10	25	4.644	55.11	35	7.128
53.00	29.2	5.670	56.04	38.4	7.803
		(1+1) + L + V			
74.80	175.5	842	75.67	200	1354
75.00	180	910	75.84	205	1491
75.15	185	1006	76.85	235	-
75.31	190	1114	77.63	260	-
75.49	195	1230			
		(6+1) + (4+1)I + V			
74.80	-15	0.27	75.31	20	3.78
75.00	0	0.92	75.49	25	5.08
75.15	+10	1.92	75.67	29.8	6.80

%	t	p	t	p	t	p
	(6+1)+(4+1)II		(4+1)I+(2+1)		(2+1)+(1+1)	
74.80	-15	0.22	-15	0.17	65	13
75.00	0	0.76	0	0.59	78	24
75.15	+10	1.62	+10	1.25	100	60
75.31	20	3.15	20	2.48	129	175
75.49	25	4.32	25	3.40	155	438
75.67	29.2	5.67	30	4.64	165	607
75.84	-	-	35	6.26	170	715
76.85	-	- -	40	8.53	175.5	842
77.63	-	-	45.3	11.77	-	-

t	p	t	p	t	p
31.2 mol %		23.8 mol %		20.4 mol %	
160	570	135	420	90	90
165	651	140	492	95	114
170	751	150	682	100	152
180	980	154	760	146	760
190	1310				
17.4 mol %		10.5 mol %		8.2 mol %	
45	30	18	5.1	-10	0.97
138	130	100	349	100	444
		120	760	114	760

Lescoeur, 1890

t	p dissoc. (2+1)	p dissoc. (6+1)	p dissoc. sat.sol.
10	-	-	3.1
15	-	-	4.3
20	-	2.3	5.4
25	-	4.0	6.75
36.5	4	-	8.5
65	13	-	32
78	24	-	57
100	60	-	134
125	175	-	-

(6+1) f.t. = 29°

Lannung, 1936

t	p dissoc.	t	p dissoc.	t	p dissoc.
0 aq.-(1+1)II		0 aq.-(1+1)I		(1+1)I-(2+1)	
0.0	0.0016	0.0	0.008	0.0	0.135
15.0	0.0092	0.0	0.007	17.4	0.540
18.0	0.0125	30.0	0.097	17.5	0.588
37.0	0.080	30.0	0.093	18.5	0.594
64.8	0.940	60.0	0.820	28.7	1.250
66.0	1.100	92.3	6.040	31.6	1.510
80.0	2.33			47.4	4.000
90.0	4.83			54.1	5.970
(2+1)-(4+1)		(4+1)II-(6+1)		(4+1)I-(6+1)	
0.0	0.244	18.0	2.01	0.0	0.60
17.2	1.25	19.9	2.41	9.7	1.36
18.0	1.33	22.0	2.90	17.2	2.55
30.0	3.65	25.0	3.81	17.5	2.61
31.7	4.28	28.0	4.96	18.0	2.73
43.4	10.24			22.0	3.78
				28.0	6.14
(2+1) sat.d.		(4+1)II sat.d.		(4+1)I sat.d.	
39.0	8.43	12.0	1.99	30.0	7.09
40.0	8.93	18.0	2.92	32.0	7.74
45.0	11.80	25.0	4.38	35.0	8.85
48.0	13.94	28.0	5.13	40.0	10.60
50.0	15.47	30.0	5.67	45.3	11.99
		34.0	6.83		
		38.0	8.04		
(6+1) sat.d.		tr.t.			
0.0	1.88	28.9	5.37	(6+1)-(4+1)II	
18.0	5.26	29.4	5.77	(6+1)-(4+1)II	
25.0	6.79	29.5	6.74	(4+1)II-(2+1)	
28.5	7.12	38.4	8.14	(4+1)I-(2+1)	
		45.3	11.99		

Robinson, 1940

Isopiestic solutions

m_1	m_2	m_1	m_2
0.0887	0.1234	0.1512	0.2127
.3117	.4534	.3692	.5453
.5481	.8476	.8242	1.361
1.027	1.781	1.105	.958
.355	2.523	.391	2.616
.583	3.098	.631	3.224
.921	4.014	.933	4.021
2.202	4.810		
0.1628	0.2304	0.1963	0.2780
.5009	.7667	.5159	.7890
.8512	1.417	.9194	1.554
1.133	2.010	1.202	2.167
.485	.843	.519	.934
.752	3.532	.765	3.580
2.107	4.534	2.178	4.747

m_1= molality of calcium chloride

m_2= molality of potassium chloride

Stokes, 1945

Isopiestic solutions

m_1	m_2	m_1	m_2
0.0908	0.1255	0.1148	0.1586
.3875	.5617	.4968	.7366
.8382	1.325	.8503	1.350
1.307	2.244	1.482	2.611
.535	.724	2.140	4.084
2.077	3.937	.921	6.008
.676	5.384		
.981	6.166		
0.2156	0.3034	0.2742	0.3901
.6090	.9232	.6808	1.046
1.028	1.683	1.170	.963
.483	2.612	.487	2.622
.772	3.241	.798	3.302
2.467	4.865	2.658	5.342
.928	6.028	.978	6.147

m_1= molality of calcium chloride

m_2= molality of sodium chloride

Stokes, 1945

Isopiestic solutions

m_1	m_2	m_1	m_2
2.951	4.329	3.072	4.520
3.929	5.899	4.126	6.237
4.635	7.116	4.928	7.611
5.473	8.557	5.511	8.622
6.394	10.091	6.425	10.136
6.862	10.780	6.874	10.813
7.286	11.383	7.326	11.434
7.430	11.582	7.431	11.575
7.914	12.199	8.023	12.331
8.693	13.458	9.785	14.371
10.750	15.429	10.771	15.442
3.363	4.985	3.715	5.561
4.422	6.742	4.525	6.916
4.967	7.687	5.366	8.381
5.611	8.801	6.064	9.547
6.640	10.456	6.861	10.786
7.092	11.115	7.233	11.315
7.341	11.456	7.354	11.483
7.525	11.715	7.775	12.041
8.193	12.552	8.749	13.209
10.071	14.701	10.159	14.800

m_1= molality of calcium chloride

m_2= molality of sulfuric acid

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.854	3.0	1.779
0.2	0.862	3.5	1.981
0.3	0.876	4.0	2.182
0.4	0.894	4.5	2.283
0.5	0.917°	5.0	2.574
0.6	0.940	5.5	2.743
0.7	0.963	6.0	2.891
0.8	0.988	6.5	3.003
0.9	1.017	7.0	3.081
1.0	1.046	7.5	3.127
1.2	1.107	8.0	3.151
1.4	1.171	8.5	3.165
1.6	1.237	9.0	3.171
1.8	1.305	9.5	3.171
2.0	1.376	10.0	3.169
2.5	1.568		

Hall, Wishow and Stokes, 1953

M	D	M	D
	25°		
0.0152	1.155	1.2965	1.236
0.0255	1.143	1.5962	1.275
0.0710	1.113	1.9765	1.306
0.1516	1.109	2.452	1.307
0.3036	1.118	2.748	1.288
0.4994	1.140	3.248	1.234
0.6426	1.160	3.530	1.192
0.9989	1.203		

Lyons and Riley, 1954

M	D	M	D
	25°		
0.02813	1.153	1.462	1.271
.0547	.136	2.046	.310
.1020	.122	2.570	.311
.1930	.123	3.250	.248
.3142	.132	4.001	.078
.4694	.152	4.486	0.919
.6706	.177	5.012	.7163
1.000	.220	5.424	.5715
.442	.271	6.004	.4020

Legrand, 1835

%	b.t.	%	b.t.
9.09	101	53.95	134
14.16	102	55.25	136
17.76	103	56.50	138
20.54	104	57.68	140
22.72	105	58.81	142
24.58	106	59.90	144
26.15	107	60.96	146
27.80	108	61.99	148
29.23	109	63.03	150
30.55	110	64.04	152
31.88	111	65.03	154
33.20	112	66.02	156
34.47	113	67.00	158
35.73	114	67.95	160
36.95	115	68.91	162
38.12	116	69.84	164
39.24	117	70.74	166
40.34	118	71.65	168
	119	72.54	170
41.38	120	73.41	172
42.40	122	74.28	174
44.38	124	75.08	176
44.44	126	75.89	178
47.97	128	76.47	179.5 (satd.)
49.59	130		
51.12	132		
52.58			

Kremers, 1856 and 1857

sat.sol. b.t.= 180°

De Heen, 1881

%	b.t.
19.02	105.8
32.65	113.0

Gerlach, 1886

%	b.t.	%	b.t.
0	100	45.80	125
5.66	101	50.24	130
10.31	102	50.64	130.4
14.16	103	54.34	135
17.35	104	57.90	140
20.00	105	61.08	145
22.48	106	64.03	150
24.53	107	66.67	155
26.20	108	68.95	160
27.80	109	71.01	165
29.33	110	72.83	170
35.48	115	74.48	175
40.83	120	75.31	178

Johnston, 1906

%	b.t.	%	b.t.
2.86	100.367	45.52	125.8
10.86	101.694	50.56	130.8
13.13	102.649	54.31	135.9
25.16	107.121	59.81	144.8
32.81	112.483	60.94	147.8
36.40	116.400	62.19	149.8
41.33	121.400	63.79	153.0

Rüdorff, 1861

%	f.t.	%	f.t.
0.99	-0.4	7.41	-3.9
1.96	-0.9	9.09	-4.9
3.85	-1.85	12.28	-7.4
5.66	-2.85	15.25	-10.0

Mulder, 1864

%	f.t.	%	f.t.
33.15	0	51.68	39
36.61	7	52.55	40
39.21	14	53.39	44.25
45.29	25.5	56.56	62
45.11	26	57.61	70
49.75	37.75	59.10	85
50.32	38.25	60.59	99

Guthrie, 1876

%	f.t.	%	f.t.
1	-0.2	10	-5.5
2	-0.5	15	-10.5
3	-1.1	20	-17.5
4	-1.6	28	-27.5 E
5	-2.1	36.45	-37 (x+1)
7	-3.3	39	0 (6+1)

Pickering, 1894

%	f.t.	%	f.t.
53	+28.55	26	-32.80
52	+29.20	25	-29.90
51	+29.43	24	-27.27
50	+29.34	23	-24.77
49	+29.02	22	-22.57
48	+28.38	21	-20.55
47	+27.46	20	-18.57
46	+26.16	19	-16.77
45	+24.62	18	-15.22
44	+22.76	17	-13.60
43	+20.40	16	-12.22
42	+17.60	15	-10.96
41	+14.47	14	- 9.78
40	+10.88	13	- 8.70
39	+ 6.72	12	- 7.72
38	+ 1.45	11	- 6.78
37	- 4.00	10	- 5.89
36	- 9.85	9	- 5.06
35	-16.25	8	- 4.31
34	-23.25	7	- 3.64
33	-31.77	6	- 3.03
32	-39.60	5	- 2.44
31	-46.95	4	- 1.91
30	-48.00	3	- 1.43
29	-43.55	2	- 0.95
28	-39.65	1.5	- 0.71
27	-35.90	1.0	- 0.46

Etard, 1894

%	f.t.	%	f.t.
31.5	-22	46.1	29
32.4	-17	49.0	35
35.1	-5	55.1	49
35.2	-5	55.9	63
36.5	+4	57.5	80
37.9	8	58.5	104
42.1	22	58.6	115

Jones and Chambers, 1900

M	f.t.	M	f.t.
0.102	-0.508	0.306	-1.537
.153	-0.752	.408	-2.104
.204	-1.012	.510	-2.681
.255	-1.267	.612	-3.348

Lidbury, 1902

mol%	f.t.	mol%	f.t.
15.29	29.312	14.31	29.918
15.11	29.492	14.25	29.920
14.95	29.637	14.15	29.913
14.82	29.714	14.07	29.904
14.76	29.758	13.97	29.854
14.68	29.800	13.85	29.800
14.62	29.833	13.74	29.734
14.53	29.866	13.59	29.652
14.45	29.890	13.44	29.530
14.38	29.911		

mol%	v.c.* 1	v.c.* 2
1.47	-	0.88
.44	1.01	1.05
.42	1.13	.17
.38	-	.28
.36	1.30	.36
.33	1.47	.60

*v.c.= velocity of crystallization, 10° below f.t.
1-diameter of tube=2.5mm
2-diameter of tube=0.5mm

Jones and Getman, 1902

M	f.t.
0.5	-2.736
1.0	6.338
1.5	11.304
2.0	17.904

Tammann, 1903

P Kg	f.t.	P Kg	f.t.
	(6+1)		
1	29.70	1953	50.15
543	35.27	2478	55.21
1003	40.17	3030	60.22
1466	45.16		

Jones and Getman, 1904; Jones and Bassett, 1904

M	f.t.	M	f.t.
0.102	-0.505	1.000	-6.345
.153	-0.752	1.500	-11.296
.204	-1.012	1.949	-17.710
.255	-1.537	2.000	-17.867
.306	-2.104	2.274	-23.000
.408	-2.681	2.598	-29.000
.510	-3.348	2.923	-37.400
.612		3.248	-46.500

Johnston, 1906

%	f.t.	%	f.t.
0.48	-2.80 ?	19.80	-16.60
2.72	-1.16	27.43	-30.00
5.48	-2.58	34.17	-52.80
10.79	-6.32		

Jones and Stine, 1908

M	f.t.	M	f.t.
0.3	-1.517	1.4	-10.05
.5	-2.672	.55	-11.58
.6	-3.270	.755	-14.33
.7	-4.065	.88	-14.69
.9	-5.422	2.2	-21.07
1.0	-6.410	2.7	-30.25
.1	-7.080	3.1	-39.50
.35	-9.362	3.5106	-49.50
0.15	-0.74	0.85	-5.05
.25	1.265	1.05	6.65
.35	1.801	.25	8.41
.45	2.35	.45	10.51
.65	3.55		

Rodebush, 1918

%	f.t.	%	f.t.
7.84	-4.21	25.86	-29.83
12.89	8.63	26.49	31.25
18.04	15.23	32.43	51.00 E

Bassett, Barton and al., 1933

f.t.	% (4+1) I	II	III
21.05	47.32	-	-
25.00	48.76	51.30	52.13
30.10	49.87	-	-
30.25	-	52.52	-
30.70	-	52.67	-
37.20	-	54.30	54.84
38.90	-	55.38	-
39.00	52.90	-	-

Yanatieva, 1946

%	f.t.	%	f.t.
15.82	-11.2	31.29	-45.5
19.14	-16.5	31.91	-39.5
22.70	-23.2	32.41	-36.2
28.10	-40.7	36.00	-8.5
30.00	-47.7 E	39.36	+7.5
30.22	-49.8(6+1)	42.07	18.5
30.48	-49.7	43.48	22.0
30.82	-49.2		

Mun and Darer, 1956

m	f.t.	m	f.t.
0.64	-3.20	1.80	-13.20
1.11	-6.60	2.46	-20.60

Kuznetsov, 1909

(4+1) I and (4+1) II

Properties of phases . Density .

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0147	18	1.1562
4	.0313	20	.1752
6	.0481	24	.2145
8	.0653	28	.2551
10	.0828	30	.2759
12	.1006	35	.3293
14	.1188	40	.3848
16	.1373	45	.4426

Grassi, 1851

%	t	d
23.049	17.5	1.218
40.99	15.8	1.417
40.99	41.25	1.339

Kremers, 1856 and 1858

%	d	%	d
	19.5°		
6.516	1.0528	26.648	1.2448
10.075	.0844	32.421	.3072
11.174	.0936	33.630	.3212
18.600	.1633	38.232	.3737
18.917	.1662	38.613	.3783
26.497	.2426		

t	d				
19.5°	1.0834	1.1632	1.2426	1.3072	1.3737
0°	1.0884	1.1705	1.2522	1.3180	-
40	.0744	.1541	.2325	.2956	1.3608
60	.0662	.1441	.2215	.2841	.3483
80	.0551	.1330	.2100	.2722	.3356
100	.0426	.1209	.1981	.2599	.3228

Schiff, 1858 and 1859

%	d	%	d	%	d
		18.3°			
0.51	1.0025	12.17	1.1046	24.33	1.2145
1.01	.0064	12.67	.1091	24.83	.2194
1.52	.0104	13.18	.1137	25.34	.2244
2.03	.0144	13.68	.1183	25.85	.2294
2.53	.0226	14.19	.1230	26.35	.2345
3.04	.0267	15.20	.1276	26.86	.2396
3.55	.0308	15.71	.1323	27.37	.2447
4.05	.0350	16.21	.1370	27.87	.2498
4.56	.0392	16.72	.1417	28.38	.2549
5.07	.0434	17.23	.1464	28.89	.2600
5.58	.0476	17.73	.1510	29.39	.2651
6.08	.0519	18.24	.1558	29.90	.2703
6.59	.0552	18.81	.1605	30.41	.2755
7.10	.0604	19.16	.1654	30.92	.2806
7.60	.0648	19.77	.1702	31.42	.2858
8.11	.0691	20.27	.1751	31.93	.2910
8.62	.0735	20.78	.1799	32.44	.2962
9.12	.0779	21.28	.1848	32.95	.3015
9.63	.0822	21.79	.1897	33.46	.3068
10.14	.0866	22.30	.1940	33.97	.3121
10.65	.0911	22.80	.1995	34.47	.3174
11.15	.0956	23.31	.2045	34.98	.3227
11.66	.1001	23.82	.2095	35.98	.3281

Gerlach, 1859

%	d	%	d	%	d
		15°			
0	0.99913	14	1.12329	28	1.26508
1	1.00764	15	.13261	29	.27592
2	.01615	16	.14232	30	.28676
3	.02466	17	.15204	31	.29803
4	.03317	18	.16175	32	.30930
5	.04268	19	.17148	33	.32058
6	.05054	20	.18119	34	.33185
7	.05940	21	.19147	35	.34312
8	.06828	22	.20174	36	.35492
9	.07714	23	.21202	37	.36671
10	.08600	24	.22229	38	.37850
11	.09532	25	.23257	39	.39029
12	.10465	26	.24341	40	.40208
13	.11397	27	.25425		

Fouqué, 1867

%	d	
	0°	11°
1.33	1.0112	1.0108
3.94	1.0378	-
24.43	1.2297	1.2262

van der Willigen, 1869

%	d	%	d
	24.0°		
16.75	1.14348	31.79	1.29697
24.38	1.22407	40.64	1.39945

Favre and Valson, 1874

N	d	N	d
	24.3°		
0	0.997	4	1.159
1	1.043	5	1.193
2	1.084	6	1.227
3	1.122	7	1.258

Kohlrausch and Grotrian, 1875 and Kohlrausch, 1879

%	d	%	d
	18°		
5	1.0409	25	1.2306
10	.0853	30	.2843
15	.1312	35	.3420
20	.1795		

Rönnberg, 1880

t	d				
	5 %	10 %	15 %	20 %	25 %
3.7	-	-	1.11565	-	-
4.0	1.04087	1.07923	-	-	1.18329
4.2	-	-	-	1.15014	-
10.0	-	-	1.11423	1.14858	1.18155
10.2	1.04014	1.07811	-	-	-
15.5	1.03915	1.07684	1.11269	1.14690	1.17965
21.0	1.03785	1.07530	1.11093	1.14500	1.17752
25.0	-	-	1.10929	-	-
25.2	-	-	-	1.14339	-
25.4	1.03660	1.07391	-	-	-
26.0	-	-	-	-	1.17542

Volkmann, 1882

%	d	%	d
	15–16°		
0	0.999	24.41	1.2269
6.50	1.0552	34.96	1.3430
12.94	1.1136		
	19°		
0	0.9984	19.99	1.1789
2.35	1.0179	29.43	1.2773
6.54	1.0540	35.98	1.3511
11.34	1.0971		

Schumann, 1887

%	d	%	d
	15°		
0	0.9991	22.50	1.2070
3.86	1.0320	23.06	.2126
8.01	.0684	28.23	.2676
14.08	.1239	37.89	.3773

Bremer, 1888

4.25%

t	d
15.65	1.033357
20.11	.032732
28.60	.029886
33.40	.028563
39.25	.026397
46.01	.023478
52.76	.020511
63.23	.015162

6.97 %

t	d
13.18	1.056608
20.97	.054529
26.48	.052915
33.27	.050526
38.45	.048494
44.70	.045832
49.93	.043546
54.93	.041199
60.02	.038550

11.97%

t	d
15.81	1.102217
24.10	.099322
35.24	.095042
44.02	.091310
50.77	.088108
56.47	.085372
73.08	.076642
100.24	.060046

22.21%

t	d
15.18	1.199718
17.97	.198524
23.37	.196129
28.38	.193903
41.50	.187647
50.76	.183014
62.93	.176663
72.08	.171552
99.80	.155167

6.96%

t	d
17.13	1.056480
25.24	.054079
29.19	.052836
35.84	.050365
42.82	.047510
48.24	.045133
58.86	.040010
66.75	.036051
75.79	.030978
99.73	.015712

10.61 %

t	d
14.16	1.086879
21.51	.084560
27.60	.082435
32.40	.080558
37.80	.078420
44.41	.075565
49.75	.073093
50.66	.072522
57.90	.069045
58.03	.069005
63.26	.066278
68.63	.063423

17.30%

t	d
6.79	1.151202
17.76	.146845
22.47	.144908
28.85	.142294
34.96	.139447
41.11	.136533
53.96	.130187
62.53	.125672
68.24	.122677

23.97%

t	d
14.57	1.213751
19.99	.211260
28.16	.207461
32.74	.205223
38.73	.202272
45.22	.198978
52.97	.194954
61.28	.190451

Drecker, 1888

%	d 0°	t	d
5.8	1.05055	22.8	1.0457
9.9	.07305	13.08	.0706
17.8	.15820	23.3	.1500
24.1	.22214	19.5	.2135
30.2	.28457	23.3	.2725
35.4	.33886	25.0	.3245
40.9	.40076	22.0	.3867

Jahn, 1891

c	d
	20°
0	0.9982
10.850	1.0830
19.912	1.1501

Charpy, 1893

c	d 0°	d 2.5°	d 53.4°
0	0.9999	0.9995	0.9865
6.5550	1.0579	1.0559	1.0552
12.4041	.1122	.1088	.1071
17.5602	.1624	.1583	.1554
22.5774	.2134	.2085	.2048
27.0563	.2614	.2556	.2511
31.2430	.3081	.3016	.2962

Pickering, 1894

17.925°

%	d	%	d
51	1.51778	22	1.19901
50	.50676	21	.18897
49	.49573	20	.17910
48	.48450	19	.16920
47	.47329	18	.15926
46	.46238	17	.14969
45	.45124	16	.14016
44	.44007	15	.13067
43	.42882	14	.12130
42	.41770	13	.11206
41	.40641	12	.10288
40	.39489	11	.09373
39	.38400	10	.08467
38	.37242	9	.07569
37	.36100	8	.06680
36	.34956	7	.05822
35	.33821	6	.04951
34	.32689	5	.04089
33	.31562	4	.03238
32	.30461	3	.02386
31	.29360	2	.01548
30	.28271	1.5	.01127
29	.27182	1.0	.00703
28	.26092	0.8	.00539
27	.25030	0.6	.00371
26	.23969	0.4	.00201
25	.22941	0.3	.00116
24	.21918	0.2	.00037
23	.20901	0.1	0.99954
		0	.99869

Sentis, 1897

mol%	t	d
1.04	17.7	1.0494
2.08	17.9	.0994
3.12	19.0	.1431
4.25	19.1	.1912
8.35	20.0	.3461

Grabowsky, 1904

%	d 10°	d 18°	d 30°
0	0.999731	0.99863	0.99567
3.56	1.028200	1.02600	1.02280
6.05	.052800	.05020	.04640
12.34	.109700	.10650	.10170
22.57	.209000	.20520	.19960
25.90	.243900	.23970	.23340
31.90	.311600	.30600	.22840

Chéneveau, 1907

%	d	%	d
16°		22°	
0	0.9990	0	0.9978
9.90	1.0813	4.97	1.0419
24.12	.2085	9.57	.0824
34.99	.3248	11.75	.1022
40.81	.3880	13.85	.1220
		17.89	.1589
		21.10	.1940
		21.58	.1958

%	t	d
6.85	14.9	1.0560
19.59	14.8	.1715
33.07	18.8	.3017

Jones and Stine, 1908

%	d	%	d
0°			
1.644	1.012908	11.15	1.086520
2.716	.021692	12.53	.107168
3.247	.025216	13.39	.118870
3.764	.032160	13.85	.122300
4.801	.040440	14.25	.125172
5.310	.045216	15.02	.133124
6.313	.054874	16.88	.154080
6.822	.057536	17.28	.156000
7.314	.062324	20.59	.186000
8.775	.075264	24.42	.227280
9.253	.079632	27.34	.258440
10.220	.086520	30.18	.287480
10.680	.091384		

Buchanan, 1912

m	d	m	d
19.5°			
6.627	1.421802	2.0	1.158200
.613	.421120	1.0	.082963
.6	.420492	0.5	.041994
.5	.416200	.25	.020544
.4	.413882	.125	.009541
.3	.407384	.062	.003979
.2	.402920	.031	.001087
.1	.398119	.016	0.999749
.0	.393585	.008	.999056
5.0	.340706	.004	.998706
4.0	.282389	.002	.998507
3.0	.225633	.001	.998421
		0.000	.998332

Siméon, 1914

c	%	d	c	%	d
15°			20°		
0	0	0.999	0	0	0.998
12.35	11.27	1.096	20.42	17.67	1.155
21.35	18.37	.162	32.13	25.97	.237
26.35	22.00	.198	41.59	32.91	.299
31.80	25.75	.235	47.34	35.36	.339
37.10	29.21	.270	50.42	37.18	.356
42.78	32.73	.307	54.25	39.31	.380
47.00	35.21	.335	57.50	41.01	.402
50.42	37.15	.357			
53.82	39.03	.379			
57.22	40.87	.400			
58.00	41.13	.410			

Tucker, 1915

mol%	d	mol%	d
16.77°			
0	0.9988	7.82	1.3308
0.92	1.0434	8.56	.3569
2.70	.1293	9.17	.3788
3.94	.1793	9.65	.3960
5.16	.2313	10.09	.4068
6.76	.2933	14.28	.5210

Gropp, 1915

t	d 6.96 N	d 4.95 N
0	1.2939	1.2136
18	1.2849	1.2084
48	1.2687	1.1971
78	1.2507	1.1826
100	1.2364	1.1656
108	1.2292	-

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
10.322	1.0786
10.593	.0806
22.087	.1665
22.665	.1675

Koch, 1924

t	d	t	d
6.25%		13.82%	
-2.2	1.0553	-6.9	1.1271
-1.8	.0552	-5.8	.1269
+1.0	.0550	-1.7	.1258
3.9	.0546	+3.5	.1245
9.7	.0536	8.2	.1231
16.1	.0522	14.1	.1211
16.2	.0521	18.2	.1197
20.3	.0510	20.1	.0091
21.43%		25.88%	
-9.7	1.2047	-29.2	1.2606
-8.8	.2043	-24.5	.2586
-3.2	.2023	-16.8	.2550
+3.1	.1997	+4.0	.2455
9.5	.1972	9.5	.2429
16.0	.1944	15.0	.2400
29.12%		37.87%	
-26.1	1.2961	-0.2	1.3821
-18.6	.2921	+0.5	.3816
-10.8	.2882	4.0	.3793
+0.4	.2824	9.5	.3758
12.7	.2760	18.2	.3702
16.5	.2741		
20.1	.2722		

Crowe, 1927

%	d					
	10°	15°	17.9°	20°	25°	30°
29.60	1.2822	1.2798	1.2783	1.2773	1.2746	1.2716
37.30	.3688	.3662	.3562	.3633	.3599	.3562
38.36	.3812	.3785	.3768	.3756	.3722	.3686
40.73	.4090	.4060	.4041	.4028	.3993	.3957
41.98	.4224	.4202	.4188	.4177	.4149	.4116
45.95	.4667	.4639	.4624	.4612	.4582	.4549

Harrison and Perman, 1927

%	d	%	d
40°			
2.26	1.014	20.52	1.196
2.92	.018	23.15	.220
4.01	.031	27.66	.265
6.69	.050	32.79	.319
10.01	.092	38.5	.383
12.57	.116	44.7	.452
14.66	.133	50.9	.520
16.27	.154	65.3	.580
19.10	.179		
50°			
3.10	1.016	22.66	1.204
5.50	.040	28.32	.252
7.58	.060	29.01	.256
9.01	.073	33.90	.300
11.59	.096	36.60	.329
12.72	.104	41.50	.380
13.94	.117	46.40	.438
16.24	.139	51.40	.500
19.00	.164	56.30	.564
68°			
1.94	0.997	23.70	1.188
2.25	0.999	26.24	.215
3.18	1.006	28.10	.244
5.51	.026	30.40	.258
6.48	.034	32.40	.280
6.69	.038	36.45	.321
7.38	.042	43.50	.395
10.36	.067	47.66	.439
12.38	.084	51.86	.474
19.33	.104	56.10	.529
22.09	.171		
70°			
3.39	1.004	27.22	1.225
5.75	.024	30.36	.260
7.61	.040	35.54	.316
8.08	.044	37.18	.334
10.11	.062	40.86	.385
12.92	.087	44.10	.407
15.85	.116	48.36	.455
19.81	.156	52.19	.494
24.13	.195	56.82	.546
25.74	.200	57.92	.556
80°			
2.48	0.994	22.55	1.162
2.93	0.998	24.07	.178
3.61	1.003	27.82	.218
5.73	.020	31.30	.257
5.87	.022	37.08	.320
8.15	.042	40.81	.353
11.60	.070	44.48	.400
14.13	.090	49.70	.460
15.40	.102	52.96	.496
18.17	.124	57.32	.546
20.20	.142	60.52	.583

Perman and Urry, 1929

c	d	c	d	c	d
30°		40°		50°	
4.211	1.0284	4.196	1.0246	4.17	1.0200
10.76	.0804	10.73	.0766	10.68	.0723
19.05	.1409	19.00	.1375	18.93	.3336
23.78	.1760	23.70	.1720	23.62	.1679
30.85	.2268	30.73	.2222	30.62	.2175
38.95	.2819	38.76	.2768	38.63	.2729
45.54	.3261	45.33	.3200	45.20	.3161
57.33	.4014	57.03	.3937	56.87	.3898
58.56	.4087	58.22	.4012	58.02	.3962
68.90	.4738	68.39	.4630	68.05	.4559
60°		70°		80°	
4.159	1.0155	4.14	1.0111	4.115	1.0049
10.64	.0678	10.59	.0632	10.52	.0559
18.84	.1282	18.75	.1229	18.63	.1157
23.51	.1622	23.38	.1564	23.24	.1491
30.45	.2110	30.30	.2050	30.16	.1995
38.41	.2656	38.24	.2585	38.03	.2533
44.92	.3083	44.71	.3019	44.53	.2967
56.57	.3822	56.30	.3758	56.02	.3686
57.77	.3895	67.35	.4407	57.19	.3759
67.72	.4486			66.92	.4322

Perman and Urry, 1929

%	c	d	%	c	d
	30°			40°	
8.43	9.993	1.0745	8.76	10.56	1.0750
15.73	22.00	.1628	15.68	21.62	.1579
21.00	33.45	.2454	20.80	32.38	.2330
25.54	45.54	.3257	25.54	45.82	.3244
28.41	55.46	.3895	28.37	54.79	.3818
30.00	61.34	.4262	31.33	65.91	.4474
	50°			60°	
8.76	10.52	1.0705	7.36	13.49	1.0491
15.68	21.52	.1528	13.54	16.17	.1201
20.81	32.25	.2277	20.70	31.92	.2200
25.54	45.61	.3179	24.81	42.78	.2942
28.37	54.54	.3750	31.77	63.27	.4221
31.33	65.60	.4406	33.12	73.60	.4808
	70°			80°	
7.36	8.308	1.0439	10.46	2.34	1.0696
13.54	16.10	.1149	15.74	21.41	.1367
20.70	31.78	.2144	20.70	31.61	.2080
24.81	42.58	.2883	23.57	38.95	.2596
31.77	62.95	.4158	26.12	46.41	.3109
33.12	73.38	.4743	29.96	59.69	.3901

Pesce, 1934

N	d	N	d
	25°		
1.706	1.07185	6.751	1.27260
2.787	.10448	8.438	.33484
3.913	.11413	9.036	.35035
4.752	.19572	10.939	.42098

Okazaki, 1935

%	d	%	d
	28°		
7.20	1.0564	20.16	1.1754
12.34	.1017	29.36	.2706
17.73	.1518	34.10	.3228

Cupples, 1945

M	m	d	M	m	d
		25°			
0.202	0.203	1.0152	1.323	1.372	1.1107
.394	.399	.0322	1.851	1.952	.1537
.589	.599	.0491	2.809	3.062	.2294
.784	.801	.0657	3.995	4.566	.3183
.966	.992	.0810	5.823	7.301	.4437

Guillaume, 1946

%	d	%	d
	20°		
6.64	1.0561	26.49	1.2460
7.09	.0570	33.90	.3169
7.35	.0640	39.10	.3763
7.66	.0670	40.00	.3980
18.00	.1594	44.00	.4285
21.66	.1965		

Rutskov, 1948

N	d		
	25°	50°	75°
5	-	1.5600	1.5395
6	1.5221	.5026	.4815
8	.4345	.4175	.3975
200	.0217	-	-
400	.00955	.0004	.9873
1000	.0021	0.9931	0.9798

Lyons and Riley, 1954

M	d	M	d
	25°		
2.518	1.2072	5.042	1.3934
3.949	.3166	5.371	.4139
4.054	.3239	5.477	.4213
4.467	.3531	5.964	.4518
4.536	.3582	6.043	.4577
4.984	.3891		

De Heen, 1881

t	r.v.	t	r.v.	t	r.v.
7.36%		32.65%		19.02%	
10.00	1.000000	10.00	1.000000	10.00	1.000000
17.40	.001959	16.30	.002749	15.90	.001945
20.20	.002867	23.22	.005650	22.32	.004184
25.51	.004610	28.70	.008011	28.25	.006368
32.31	.007156	34.57	.010584	35.56	.009189
40.67	.010523	40.52	.013311	42.22	.011955
50.14	.014938	47.17	.016152	49.03	.014844
		54.42	.019333	59.75	.019561
				65.27	.022235

r.v.= relative volume

Zepernick and Tammann, 1895

t	relative volume 2.781%	5.644%	10.577%	27.97%
0	1	1	1	1
110	1.0517	1.0509	1.0510	1.0500
120	.0597	.0588	.0585	.0568
130	.0688	.0676	.0663	.0626
140	.0784	.0764	.0744	.0692

Grassi, 1851

%	t	π
0	13.4	47.7
0	18.0	46.2
0	43.0	44.2
23.049	17.5	30.6
40.99	15.8	20.6
40.99	41.25	22.9

Drecker, 1888

%	π
0°	
0	44.1
5.8	39.7
9.9	37.1
17.8	31.3
24.1	27.6
30.2	25.6
35.4	23.2
40.9	21.7

Perman and Urry, 1929

c	π 0-100°	π 100-200°	c	π 0-100°	π 100-200°
4.211	41.4	39.9	4.196	40.9	39.8
10.76	37.6	36.1	10.73	37.3	36.0
19.05	33.3	32.0	19.00	33.2	32.0
23.78	31.2	30.1	23.70	31.2	30.1
30.85	28.8	28.0	30.73	28.8	28.0
38.95	26.6	25.9	38.76	26.6	25.9
45.54	24.9	24.3	45.33	25.1	24.4
57.33	22.8	22.3	57.03	23.0	22.3
58.56	22.5	22.0	58.22	22.6	22.1
68.90	21.2	20.8	68.39	21.4	20.9
4.17	40.7	39.8	4.159	41.2	40.0
10.68	37.2	35.9	10.64	37.6	36.3
18.93	33.3	32.1	18.84	33.5	32.4
23.62	31.2	30.2	23.51	31.5	30.5
30.62	28.8	28.1	30.45	29.0	28.3
38.63	26.6	26.0	38.41	27.0	26.1
45.20	25.1	24.5	44.92	25.3	24.7
56.87	23.0	22.4	56.57	23.3	22.8
58.02	22.8	22.2	57.77	22.9	22.5
68.05	21.6	21.1	67.72	21.8	21.3
4.14	41.5	40.3	4.115	42.3	41.0
10.59	37.9	36.5	10.52	38.6	37.2
18.75	33.7	32.7	18.63	34.2	33.2
23.38	31.6	30.7	23.24	32.1	31.1
30.30	29.4	28.5	30.16	29.7	28.9
38.21	27.3	26.3	38.03	27.6	26.6
44.71	25.5	24.9	44.53	25.8	25.2
56.30	23.5	22.8	56.02	23.8	23.1
67.35	22.0	21.5	57.19	23.6	23.0
			66.92	22.2	21.8

Viscosity and surface tension .

Grotrian, 1876

%	t	η	t	η
5.00	7.92	2187	15.91	1686
9.98	7.29	2766	15.66	2042
19.93	7.93	3943	15.92	3115
25.38	7.78	5525	14.34	4448
29.81	8.04	7993	16.57	5976
35.2	8.13	15003	16.35	11130

%	η	$\tau . 10^4$	%	η	$\tau . 10^4$
18°					
5.00	1555	264	25.38	3847	257
9.98	1840	282	29.81	5665	277
19.93	2899	245	35.20	10350	293

Hechler, 1904

t	η	t	η	t	η
30.283%		32.222%		32.533%	
-20.42	7545	-20.65	8889	-20.63	9190
20.10	7450	20.58	8803	20.23	8975
20.03	7428	20.22	8711	19.72	8825
19.45	7279	19.29	8292	19.01	8613
18.19	6787	18.83	8263	18.16	8320
17.06	6494	17.71	7917	17.05	7980
15.41	6057	17.08	7611	14.31	7107
13.93	5712	16.03	7429	13.42	6877
12.76	5430	14.76	7006	11.56	6410
11.66	5210	13.56	6698	10.57	6145
10.72	5050	12.27	6390	-9.67	5958
9.72	4830	11.38	6120	+0.09	4240
-8.41	4590	-6.85	5220	0.17	4240
+0.05	3520	+0.21	4122		
7.00	2907	5.49	3500		
7.58	2860	7.92	3250		
9.27	2740	9.71	3060		
11.27	2630	11.09	2960		
12.26	2540				

Siméon, 1914

c	%	η	c	%	η
	15°			20°	
0	0	1133.6	0	0	997.4
12.35	11.27	1528	20.42	17.67	1704
21.35	18.37	1980	32.13	25.97	2542
26.35	22.00	2324	41.59	32.99	3817
31.8	25.75	2795	47.34	35.36	4946
37.1	29.21	3468	50.42	37.18	6143
42.78	32.73	4604	54.25	39.31	7603
47.0	35.21	5861	57.50	41.01	9733
50.42	37.15	7089			
53.82	39.03	8756			
57.22	40.87	10945			
58.00	41.13	11709			

Walker, 1914

d	η	d	η
17.2°			
0.998	1079	1.113	1522
1.0170	1132	.128	1591
.052	1241	.147	1767
.061	1277	.163	1824
.082	1370	.182	2006
.1009	1450	.2057	2429

Tucker, 1915

mol%	η	mol%	η
16.77°			
0	1081	7.82	4860
0.92	1210	8.86	6020
2.70	1610	9.17	7060
3.94	1990	9.65	8180
5.16	2510	10.09	9890
6.76	3710	14.28	33200

Stakelbeck and Plana, 1929

%	η 20°	10°	0°	-10°
0.8	1025	1335	1825	-
1.6	1035	1345	1840	-
2.5	1055	1365	1860	-
3.4	1075	1385	1890	-
4.2	1100	1405	1925	-
5.0	1125	1435	1960	-
5.9	1150	1460	1995	-
6.8	1175	1490	2040	-
7.8	1205	1520	2090	-
8.7	1240	1555	2145	-
9.6	1275	1595	2200	-
10.5	1310	1640	2255	-
11.3	1345	1685	2315	-
12.2	1380	1735	2380	-
13.1	1425	1785	2455	-
14.1	1480	1845	2540	-
15.0	1540	1905	2630	4155
16.0	1610	1985	2730	4300
16.9	1680	2065	2835	4450
17.8	1755	2155	2948	4610
18.8	1840	2255	3055	4780
19.8	1925	2355	3180	4990
20.8	2020	2475	3325	5210
21.8	2130	2610	3480	5440
22.7	2240	2745	3645	5680
23.7	2370	2900	3835	5965
24.7	2520	3080	4050	6310
25.7	2675	3260	4300	6715
26.6	2830	3465	4600	6720
27.6	3025	3750	4950	8325
28.7	3280	4070	5350	9035
29.7	3580	4420	5800	9200
29.9	3630	4500	5900	9850
30.7	3880	4780	6285	10815
31.8	4200	5200	6870	11875
32.9	4540	5690	7560	13175
34.0	4900	6240	8325	14525
35.1	5260	6800	9180	-
36.1	5650	7375	10090	-
37.2	6050	8025	11075	-

%	η -15°	-20°	-25°
18.8	6240	-	-
19.8	6420	-	-
20.8	6690	-	-
21.8	6950	8690	-
22.7	7290	9005	-
23.7	7625	9510	11675
24.7	8050	10060	12320
25.7	8565	10720	13155
26.6	9115	11400	14000
27.6	9790	12250	15250
28.7	10525	13300	16690
29.7	11385	14550	-
29.9	11660	14845	-
30.7	12420	15850	-
31.8	13705	17600	-
32.9	15270	19300	-

Lyons and Riley, 1954

M	η	M	η
	25°		
0.0953	921.7	2.518	2023
.1918	949.4	3.202	2713
.3724	1000	.298	2842
.6373	1080	.949	3988
.7520	1115	4.054	4232
.7798	1126	.467	5399
.8797	1160	.536	5645
1.0824	1231	.984	7577
.1155	1244	5.042	7961
.398	1362	.371	9843
.525	1416	.477	10648
.917	1620	.964	15207
2.2016	1674	6.043	16339

Volkmann, 1882

%	σ	%	σ
15-16°		19°	
0	73.2	0	72.7
6.50	75.7	2.35	73.7
12.94	78.0	6.54	74.7
24.41	84.1	11.34	76.6
34.96	91.1	19.99	80.7
		29.43	86.7
		35.98	91.3

Ochsé, 1890

t	a^2				
	0 c	5 c	10 c	15 c	20 c
4	78.92	-	-	58.99	56.09
5	-	65.17	61.88	-	-
8	76.75	-	-	-	-
15	74.25	63.77	59.88	57.19	55.19
25	71.46	61.98	58.68	55.39	53.19
35	-	60.18	56.69	53.79	51.30
40	85.37	-	-	-	-
45	-	57.19	54.19	52.10	49.90
55	61.38	54.59	52.60	50.90	49.00

c = g/100 cc of water .

Sentis, 1897

mol%	t	σ	mol%	t	σ
0	25.1	72.3	3.12	19.0	79.0
0	13.5	74.0	4.25	19.1	81.5
1.04	17.7	75.1	8.35	20.0	91.7
2.08	17.9	77.0			

Grabowsky, 1904

%	σ		%	σ	
	10°	30°		10°	30°
0	75.35	72.33	22.57	85.32	81.75
3.56	76.55	73.55	25.90	87.20	84.16
6.05	77.47	74.49	31.90	91.21	87.89
12.34	79.95	76.99			

Morgan and Schramm, 1913

%	σ	%	σ
	30°		
0	71.03	40.72	88.43
2.86	71.74	41.57	90.85
7.30	72.87	44.16	91.92
9.73	73.48	45.90	92.80
12.27	74.67	47.41	94.96
15.09	75.56	48.65	95.79
17.55	76.71	49.46	97.06
19.69	78.04	50.66	98.12
21.77	78.64	51.40	99.10
23.77	80.03	51.82	99.34
27.72	82.44	52.17	99.69
28.37	82.77	53.04	100.19
30.39	83.78	53.15	100.56
33.99	85.13	54.65	101.79
34.76	87.12	54.76	102.02
37.16	87.65	55.50	102.57
37.22	87.73		

Cupples, 1945

M	m	σ	M	m	σ
		25°			
0.202	0.203	72.43	1.323	1.372	76.25
.394	.399	73.06	1.851	1.952	78.33
.589	.599	73.59	2.809	3.062	82.81
.784	.801	73.35	3.995	4.566	88.49
.966	.992	74.70	5.823	7.301	97.20

Optical and electrical properties .

Kremers, 1856 and 1857

%	n_D
16°	
0	1.3333
24.75	.3942
40.41	.4388

van der Willigen, 1869

w.l.	n_D 16.75%	24.38%	31.79%	40.64%
	24°			
A	1.36910	1.39107	1.41060	1.43722
a	.37009	.39211	.41165	.43834
B	.37090	.39296	.41257	.43929
C	.37175	.39392	.41353	.44034
D	.37392	.39633	.41611	.44313
E	.37667	.39932	.41936	.44668
b	.37718	.39987	.41999	.44733
F	.37899	.40187	.42216	.44972
G	.38135	.40447	.42502	.45287
G	.38326	.40660	.42728	.45545
H	.38405	.40745	.42825	.45655
H	.38535	.40889	.42982	.45829
H	.38699	.41059	.43179	.46035

Jahn, 1891

c	n H_α	D	H_β
	20°		
0	1.3315	1.3332	1.3375
10.850	.3552	.3572	.3619
19.912	.3738	.3761	.3811

Bremer, 1900

%	t	n H_α	D	H_β	H_γ
0	16	1.3308	1.3332	1.3371	1.3389
7.0110	20	.34914	.35089	.35491	.35866
9.3168	17	.35349	.35510	.35992	.36418
12.2954	17	.35776	.35897	.36419	.36793
11.7576	20	.36040	.36225	.36637	.36962
13.0884	16	.36405	.36631	.37100	.37474
19.6565	13	.38137	.38323	.38842	.39223

Chéneveau, 1907

%	n_D	%	n_D
	16°		
0	1.3333	34.99	1.4238
9.90	1.3563	40.81	1.4407
24.12	1.3922		
	22°		
0	1.3328	13.85	1.3666
4.97	1.3442	17.89	1.3773
9.57	1.3556	21.10	1.3877
11.75	1.3611	21.58	1.3879

%	t	n C	D	Tl	F	G'
6.85	14.9	1.34797	1.34995	1.35275	1.35439	1.36793
19.59	14.8	.38085	.38309	.38589	.38831	.40244
33.07	18.8	.41582	.41842	.42122	.42440	.42927

Jahn, 1891

c	(α)magn.
	20°
0	1
10.850	1.5311
19.912	1.4897

Schönrock, 1893

%	(α)magn.
	16°
0	1
12.1638	1.5798

Guillaume, 1946

%	(α)* magn. 10^6	n 5678 Å
	20°	
6.64	4.148	1.3501
7.09	.153	.3505
7.35	.158	.3512
7.66	.165	.3525
18.00	.405	.3794
21.66	.465	.3900
26.49	.554	.4032
33.90	.664	.4225
39.10	.713	.4370
40.00	.723	.4436
44.00	.721	.4509

* in radians, gauss, centim.

Kohlrausch and Grotrian, 1875

%	κ 0°	18°
5	410	639
10	744	1134
15	999	1495
20	1046	1718
25	1164	1770
30	1077	1638
35	843	1357

Grotrian, 1876

%	κ	$\tau.10^4$
	18°	
5.00	639	209
9.98	1132	201
19.93	1715	196
25.38	1768	200
29.81	1658	209
35.20	1341	229

Kohlrausch, 1879

%	κ	$\tau.10^4$
	18°	
5	639	214
10	1134	207
15	1495	203
20	1718	201
25	1770	205
30	1648	217
35	1357	237

Kunz, 1902

t	κ
25.52 %	
0	1159.65
-11.56	834.08
-20.69	638.92
-28.32	453.49
29.0 %	
0	1102.84
-11.87	771.19
-20.81	596.56
-30.36	368.58

Hechler, 1904

t	κ	t	κ	t	κ
30.283%		32.222%		32.533%	
-22.52	520	-20.64	517	-33.58	285
21.42	544	20.62	518	32.58	300
20.40	564	20.27	526	30.58	331
20.18	571	19.43	545	29.58	348
20.04	572	18.86	556	28.09	373
19.38	588	18.82	554	20.60	514
18.02	619	17.77	583	20.24	522
16.95	644	17.36	589	19.73	531
15.29	685	16.24	617	19.06	547
14.03	715	14.93	646	18.21	566
12.87	746	13.70	673	17.12	590
11.78	772	12.43	704	14.41	651
10.84	796	11.51	725	13.54	673
9.79	824	9.17	784	11.71	716
-8.82	852	-7.03	838	10.64	743
+0.18	1105	+0.09	1037	-9.74	765
0.14	1104	0.22	1040	+0.03	1031
6.95	1323	5.18	1186	.10	1033
7.44	1339	6.04	1213	.09	1033
9.24	1401	7.88	1270	.16	1038
11.24	1466	9.64	1325		
12.25	1500	11.07	1372		

Jones and Stine, 1908

M	κ	M	κ
	10°		
0.15	149.91	0.85	669.03
0.25	240.73	1.05	782.44
0.35	327.00	1.25	888.53
0.45	396.88	1.45	973.35
0.65	542.36		

Gropp, 1915

t	κ	t	κ
6.960 N		4.950 N	
-52.9	86.92	-56.8	48.20
40.45	218.5	48.6	88.28
29.8	389.0	46.9	103.6
24.6	486.4	37.2	242.4
19.85	588.1	26.4	547.8
-10.05	831.0	15.8	786.0
0	1080	6.2	1002
+18	1670	0	1171
48	2811	+18	1770
78	4069	48	2924
100	5028	78	4130
108	5361	100	5044

Tucker, 1915

mol%	κ	mol%	κ
16.77°			
0.92	492.1	7.82	1032
2.70	1093.0	8.66	931
3.94	1278.0	9.17	847.9
5.16	1304.0	9.65	782.3
6.76	1183.0	10.09	696.9
		14.28	310.6

Crowe, 1927

%	κ					
	10°	15°	17.9°	20°	25°	30°
29.60	1378	1556	1658	1732.	1912	2097
37.30	976	1115	1198	1261	1413	1577
38.36	907	1040	1121	1181	1325	1475
40.73	743	859	931	986	1121	1260
41.98	679	798	866	919	1044	1178
45.95	426	521	574	613	712	818

Pesce, 1934

N	κ	N	κ
25°			
1.706	1168	6.751	1471
2.787	1233	8.438	1550
3.913	1327	9.036	1644
4.752	1373	10.939	1679

Heim, 1886

t	relative resistance
41.69%	
40	111.6
35	124.0
30	138.4
25	155.6
20	176.5
15	203.4
10	234.8
5	272.6
3.5	286.1
-0.5	330.3
-1.5	342.5
-2.5	356.4
-3.5	369.7
-4.5	383.6

Mc Gregory, 1894

N	λ	N	λ
18°			
0.0001	114.67	0.1	87.29
.0002	113.34	0.5	74.43
.0006	112.04	1.0	67.3
.001	110.78	2.0	57.5
.002	110.05	3.0	49.3
.006	104.88	4.0	42.1
.01	102.56	5.0	35.4
.05	92.33	6.0	29.1

Jones and Chambers, 1900

M	molecular conductivity
25°	
0.1020	174.05
0.1530	164.75
0.2040	156.62
0.2556	152.06
0.5100	134.30

Jones and Getman, 1902

M	molecular conductivity
0°	
1.0	71.16
1.5	62.14
2.0	53.98
2.5	44.79

Jones and Getman, 1904 and Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
0°			
0.102	105.70	1:000	71.15
0.153	102.90	1.500	62.06
0.204	98.40	1.949	48.83
0.255	96.92	2.000	54.05
0.306	89.61	2.274	44.44
0.408	89.10	2.598	39.55
0.510	88.24	2.923	35.88
0.612	84.25		

Desai, Naik and Desai, 1934

N	λ
30°	
0.005	125.08
.01	117.00
.25	79.03
1.00	64.90
5.0	33.01

Okazaki, 1942

%	Verdet's constant (D).10^5
25°	
11.11	1527
15.55	1625
19.92	1718
24.42	1817
28.10	1902
32.32	1999
37.82	2122

Okazaki, 1935

%	Verdet's constant. 10^5 (3441 Å)
28°	
7.20	4908
12.34	5347
17.73	5794
20.16	6023
29.36	6863
34.10	7267

Mc Clung and Mc Intosh, 1902

d	X-ray absorption (in %)
room temperature	
1.000	63.0
.018	65.7
.034	73.0
.069	78.8
.135	88.3
.259	98.0

Thermal constants.

Marignac, 1876

%	U	%	U
21-51°			
2.98	0.9550	19.75	0.7538
5.81	.9174	29.09	.6741
10.96	.8510	38.11	.6176

Drecker, 1888

%	U	%	U
0°			
0	1.000	24.1	0.730
5.8	0.936	30.2	.685
9.9	.876	35.4	.656
17.8	.789	40.9	.636

Pickering, 1894

%	U	%	U
0.4284	0.9994	12.3944	0.8841
1.6957	.9860	22.6615	.7683
2.3450	.9707	39.060	.6253
6.5145	.9370		

Teudt, 1900

t	U	t	U	t	U
13.24%		19.93%		34.90%	
32.1	0.8126	31.3	0.7571	32.3	0.6456
33.1	.8254	31.9	.7535	32.5	.6485
33.6	.8245	37.1	.7672	35.4	.6494
37.7	.8243	37.3	.7672	36.7	.6498
38.2	.8241	39.9	.7736	38.2	.6505
43.6	.8367	39.4	.7721	42.3	.6609
43.8	.8349	41.4	.7695	43.3	.6639
50.1	.8424	44.7	.7804	50.1	.6620
50.4	.8465	47.7	.7828	50.9	.6645
53.6	.8527	51.3	.7792	51.7	.6675
		52.4	.7909		
		54.8	.7808		

Dickinson, Mueller and George, 1910

t	U			
	d^{20}=1.07	d^{20}=1.14	d^{20}=1.20	d^{20}=1.26
-25	-	-	-	0.648
-20	-	-	0.695	.651
-15	-	0.764	.700	.654
-10	-	.768	.705	.657
-5	0.873	.772	.709	.660
0	.877	.775	.712	.663
+5	.880	.778	.715	.667
+10	.882	.781	.719	.670
+15	.884	.784	.722	.673
+20	.887	.787	.725	.676

Koch, 1922

%	U				
	-40°	-30°	-20°	-10°	0°
0	-	-	-	-	1.005
2	-	-	-	-	0.976
4	-	-	-	-	.945
6	-	-	-	-	.914
8	-	-	-	-	.885
10	-	-	-	-	.858
12	-	-	-	-	.830
14	-	-	-	-	.803
16	-	-	-	0.774	.779
18	-	-	-	.751	.756
20	-	-	-	.729	.735
22	-	-	0.704	.710	.716
24	-	-	.687	.693	.699
26	-	0.665	.671	.677	.683
28	0.644	.650	.656	.662	.668
30	.630	.636	.642	.648	.654
32	-	.621	.627	.633	.639
34	-	-	-	.620	.626
36	-	-	-	-	.613

%	U				
	10°	20°	30°	40°	50°
0	1.001	0.999	0.998	0.998	0.999
2	0.975	.974	.973	.972	.971
4	.946	.946	.947	.947	.948
6	.916	.917	.919	.920	.922
8	.888	.890	.893	.895	.898
10	.860	.863	.866	.871	.872
12	.848	.838	.842	.846	.850
14	.808	.813	.818	.823	.828
16	.784	.789	.795	.800	.805
18	.762	.768	.773	.779	.784
20	.741	.746	.752	.758	.764
22	.722	.728	.734	.740	.746
24	.705	.711	.717	.723	.729
26	.689	.695	.701	.708	.714
28	.674	.680	.686	.692	.698
30	.660	.666	.672	.678	.684
32	.645	.651	.658	.665	.671
34	.633	.639	.646	.652	.659
36	.620	.627	.633	.640	.647
38	.606	.614	.621	.628	.635
40	-	.602	.609	.617	.624

Mueller, 1929

t	U			
	$d^{15.6}$=1.175	=1.200	=1.225	=1.250
-23.3	-	-	0.670	0.654
17.8	0.722	0.697	.676	.659
12.2	.728	.703	.681	.663
6.7	.733	.708	.685	.667
1.1	.736	.711	.689	.670
+4.4	.740	.715	.693	.674
10.0	.743	.719	.697	.677
15.6	.746	.722	.700	.680
21.1	.750	.726	.704	.684

Rutskov, 1948

%	U		
	25°	50°	75°
0.612	0.9890	0.9893	-
1.517	.9758	.9764	0.9801
2.000	.9686	-	-
2.989	.9544	0.9554	0.9598
5.000	.9266	-	-
5.804	.9155	0.9179	0.9228
10.000	.8610	-	-
10.97	.8495	0.8530	0.8588
14.61	.8055	.8104	-
19.77	.7480	.7548	0.7638
25.50	.6972	.7050	.7145
33.92	.6377	.6468	.6583
38.12	.6110	.6205	.6325
43.51	.5780	.5885	.600
50.66	-	.5345	.548
55.20	-	:5010	.510

Hammerl, 1879

%	t	Q mix (cal/g) (with snow)	%	t	Q mix (cal/g) (with snow)
78	+7.6	-24.50	69	-14.1	-21
75	+2.9	-24.12	67	-19.7	-37
74	0	-23.92	65	-26.4	-75
73	-2.2	-23.78	63	-33.3	-54
71	-8.4	-23.48			

M	t	Q diss. (cal/g salt)	M	t	Q diss. (cal/g salt)
78	+ 7.6	-	60	-49.5	- 74.8
75	+ 2.0	-	59	-54.9	- 77.9
74	+ 0.9	-52.2	57	-48.2	- 81.1
73	- 2.2	-53.4	55	-40.3	- 87.1
72	- 4.3	-55.1	54.5	-39.9	- 87.4
71	- 8.4	-56.4	52	-36.5	- 90.9
70	-10.6	-58.2	50.5	-30.43	- 99.0
69	-14.1	-59.5	49.5	-27.99	-103.2
68	-17.5	-61.2	46	-22.7	-116.0
67	-19.7	-62.4	45	-21.5	-119.3
66	-22.8	-63.7	42	-18.3	-131.9
65	-26.4	-65.5	38	-14.7	-151.8
64	-28.7	-66.7	35	-12.4	-171.3
63	-33.3	-68.6	29	- 9.0	-216.9
62	-39.0	-71.1	27	- 8.1	-237.3
61.5	-41.2	-71.7	23	- 6.5	-282.4
			17	- 4.0	-413.6

Pickering, 1894

%	Q dil.	%	Q dil.
	17.91°		
51.781	3285	37.983	1041
51.716	3265	37.005	962
51.759	3278	35.913	867
51.511	3232	34.796	788
51.006	3124	33.700	712
50.569	3008	32.689	653
49.774	2881	31.428	533.2
49.371	2790	30.950	562
48.784	2688	29.947	514
48.278	2581	28.371	447
47.584	2457	26.866	393.5
46.974	2335	25.834	367.4
46.278	2214	24.363	322.3
45.176	2034	23.767	306.0
44.382	1911	22.6615	236.7
43.615	1768	19.995	217.1
43.290	1718	17.518	165.9
42.914	1660	15.060	136.7
42.195	1564	12.3944	107.3
42.103	1568	6.515	47.1
42.149	1566	3.345	20.91
41.484	1464	1.696	9.04
40.679	1373	0.854	3.68
40.275	1307	0.428	1.45
39.060	1164	0.214	0.45

Q dil. of 100 g in a very large amount of H_2O .

Harrison and Perman, 1927

%	Q dil.	%	Q dil.
	40.00°		
51.85	-77.13	30.52	- 9.74
51.31	-75.39	24.94	- 4.57
50.87	-73.31	24.26	- 4.46
50.47	-71.20	19.24	- 2.20
44.56	-47.03	14.92	- 1.12
40.20	-32.38	9.20	- 0.67
35.13	-18.32		
	50.00°		
50.56	-77.64	24.31	- 5.62
47.60	-65.43	19.46	- 2.70
40.67	-35.48	15.53	- 1.61
34.47	-20.65	11.55	- 0.85
29.34	-10.66		
	60.00°		
54.03	-100.5	26.17	- 6.49
48.86	-78.21	20.20	- 2.97
42.11	-49.08	15.30	- 1.65
41.70	-46.55	11.03	- 0.60
32.85	-17.29		
	70.00°		
49.61	-71.20	24.21	- 6.65
46.77	-69.79	18.16	- 3.13
37.68	-35.07	12.28	- 1.10
31.11	-13.79		
	80.00°		
49.48	-92.51	26.68	-10.85
46.43	-73.43	23.24	- 6.41
38.20	-41.07	17.99	- 3.45
35.65	-29.71	13.10	- 1.48
33.29	-23.41		

Q dil. : addition of 1g H_2O to an infinite amount of solution .

Lehtonen, 1922

M	Q mix (cal/g)	M	Q mix (cal/g)
	0°		
0.0625	144.717	0.9998	141.088
0.1250	144.176	2.0000	138.863
0.2500	143.633	4.0000	135.289
0.5000	142.630		

Beetz, 1879

d^{20}	t	heat conductivity.10^5
1.345	8-14°	431
-	28-36°	550
1.128	8-14°	424
-	28-36°	606
water	8-14°	413
-	28-36°	662

Kapustinskii and Ruzavin, 1955

Heat conductivity coefficient k.10^6

%	k
25°	
4.50	1446
9.60	1426
14.5	1416
24.4	1387
36.6	1350

Cal in cm/sec/°C

Water + Calcium bromide ($CaBr_2$)

Heterogeneous equilibria .

Tammann, 1885

t	p 0%	18.73%	27.16%	40.33%	45.30%
42.45	63.0	59.0	54.1	43.8	37.4
47.29	80.8	75.1	69.3	65.7	47.8
51.28	98.6	91.8	85.0	68.1	58.3
55.49	120.9	112.8	104.3	83.7	72.0
61.40	159.4	148.3	139.1	111.3	95.8
64.42	182.8	170.7	159.7	127.4	110.5
67.88	213.2	199.0	186.0	149.5	132.0
70.71	241.1	224.1	209.6	168.6	145.9
73.65	273.3	254.9	249.0	191.3	165.9
77.75	324.3	302.4	284.1	228.0	198.3
80.66	365.1	340.4	319.6	257.0	224.2
83.68	411.9	384.7	361.4	290.9	253.8
88.45	495.7	462.0	434.9	352.8	310.3
91.39	554.4	518.7	486.8	393.3	344.4
94.92	632.2	590.4	554.0	448.7	393.4
96.93	680.3	634.6	596.9	484.7	427.7
100.57	775.6	723.2	680.0	555.5	504.0

%	p	%	p
	100°		
6.43	748.5	39.28	547.0
17.62	714.6	43.35	490.7
20.21	703.4	44.01	480.5
25.48	674.5	47.52	425.2
29.71	644.6	51.25	379.2
35.64	588.9	52.91	334.8

Robinson, 1942

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.872	0.9	1.097
0.2	0.881	1.0	1.133
0.3	0.900	1.2	1.208
0.4	0.926	1.4	1.285
0.5	0.958	1.6	1.365
0.6	0.992	1.8	1.447
0.7	1.027	2.0	1.531
0.8	1.062		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.863	1.6	1.370
0.2	0.878	1.8	1.455
0.3	0.900	2.0	1.547
0.4	0.927	2.5	1.790
0.5	0.958	3.0	2.048
0.6	0.990	3.5	2.297
0.7	1.022	4.0	2.584
0.8	1.057	4.5	2.908
0.9	1.093	5.0	3.239
1.0	1.131	6.0	3.880
1.2	1.207	7.0	4.463
1.4	1.286	8.0	4.809

Johnston, 1906

%	b.t.	%	b.t.
2.42	100.079	47.58	106.787
6.78	100.296	49.59	107.563
13.62	100.679	51.27	108.387
20.10	101.176	52.70	109.107
24.99	101.616	54.99	111.545
29.25	102.116	55.79	115.107
33.70	102.723	70.33	128.400
36.25	103.413	71.98	140.400
40.15	104.233	72.30	142.400
43.04	105.023	73.15	144.000
45.28	106.018	73.95	145.400

Etard, 1894

%	f.t.	%	f.t.
50.5	-22	53.1	+ 8
50.2	-22	55.1	+ 9
52.5	-14	55.7	+11
52.6	- 7	57.1	+20
52.6	- 5	62.6	+50

Jones and Getman, 1904 and Jones and Bassett, 1905

M	f.t.	M	f.t.
0.0435	-0.228	0.5226	- 2.949
0.0871	-0.445	0.452	- 2.340
0.1306	-0.664	0.903	- 6.200
0.1742	-0.904	1.506	-13.100
0.2613	-1.368	1.807	-17.500
0.3484	-1.847	2.409	-30.500
0.4355	-2.397	3.011	-47.000

Kuznetsov, 1909

(4+1) II   f.t. = 55°

Kremers, 1856 and 1859

%	d	%	d
	19.5°		
5	1.042	30	1.313
10	1.087	35	1.383
15	1.137	40	1.459
20	1.192	45	1.638
25	1.250	50	1.638

t	d					
19.5°	1.1387	1.2799	1.3995	1.5135	1.6126	1.7525
0	1.14418	1.28882	1.4109	1.5270	1.6275	1.7688
40	.13020	.2692	.3870	.4993	.5971	.7358
60	.11998	.2576	.3741	.4853	.5822	.7197
80	.10804	.2449	.3607	.4711	.5672	.7037
100	.09473	.2313	.3467	.4566	.5524	.6880

Jahn, 1891

c	d
20°	
0	0.9982
16.51	1.1337
30.94	1.2490

Heydweiller, 1909

N	d	N	d
18°			
0.05	1.0029	1.0	1.0809
.10	.0071	2.0	.1617
.20	.0154	4.0	.3180
.05	.0401		

Thomas and Perman, 1934

%	d
30°	
13.29	1.132
26.84	.295
40.36	.480
52.76	.663
60.04	.798

Okazaki, 1935

%	d	%	d
28°			
7.04	1.0574	32.06	1.3363
14.30	.1273	35.79	.3880
21.81	.2093	40.26	.4550
27.13	.2726	44.72	.5300

Guillaume, 1946

%	d	%	d
20°			
11.0	1.0993	32.5	1.3473
14.0	.1287	47.5	.5898
21.54	.2120	59.5	.8117

Thomas and Perman, 1934

%	π
30°	
13.29	38.2
26.84	32.2
40.36	27.2
52.76	22.1
60.04	20.4

Kremers, 1857

%	t	n_D
0	17	1.3332
38.90	18	.4188
55.88	17	.4776

Guillaume, 1946

%	(α)magn.10^6*	n 5780 Å
20°		
11.0	4.250	1.3541
14.0	.310	.3600
21.54	.487	.3771
32.5	.729	.4041
47.5	5.025	.4501
59.5	5.179	.4906

*in radians, gauss, centim.

Jahn, 1891

c	n H_α	n D	n H_β	(α)magn.
20°				
0	1.3315	1.3332	1.3375	-
16.51	.3583	.3604	.3654	1.6064
30.94	.3812	.3836	.3893	1.5643

Jones and Chambers, 1900

M	λ
25°	
0.174	87.18
.218	85.44
.456	75.60

Jones, 1904; Jones and Getman, 1904 and Jones and Bassett, 1905

M	λ	M	λ
	0°		
0.0435	53.71	0.5226	42.93
.0871	50.26	0.903	40.45
.1306	49.14	1.506	34.15
.1742	47.26	1.807	30.85
.2613	46.59	2.409	25.07
.3484	44.72	3.011	19.84
.4355	43.79		

Johnston, 1906

N	λ	N	λ
	99.4°		
0.001	382.1	1	150.1
.01	311.6	2	140.2
.10	202.8	4	103.8
.25	185.8	6	76.9
.50	168.5	8	59.0
		10	41.0

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	98.0	1.0	71.8
.10	92.7	2.0	62.0
.20	87.1	4.0	45.7
.50	79.0		

Okazaki, 1935

%	Verdet's constant (3441 Å)
28°	
7.04	0.05004
14.30	.05714
21.81	.06492
27.13	.07149
32.06	.07805
35.79	.08313
40.26	.08996
44.72	.09679

Water + Calcium iodide (CaI_2)

Dieterici, 1898

M	p
0°	
0	4.579
0.992	4.417
2.032	4.122
4.845	3.033
6.934	2.285

Stokes, 1948

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.880	0.9	1.173
.2	.906	1.0	.217
.3	.935	1.2	.310
.4	.969	1.4	.407
.5	1.008	1.6	.504
.6	.044	1.8	.605
.7	.083	2.0	.710
.8	.128		

Robinson, 1942

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.883	0.9	1.177
.2	.909	1.0	.222
.3	.939	.2	.313
.4	.973	.4	.409
.5	1.007	.6	.507
.6	.046	.8	.613
.7	.088		
.8	.132		

Etard, 1894

%	f.t.	%	f.t.
61.6	-22	69.4	51
65.0	+7	75.9	64
65.1	+10	81.3	130
66.3	19	87.1	248

Jones and Getman, 1904

M	f.t.	M	f.t.
0.078	-0.374	1.248	-10.030
.156	0.743	2.184	27.000
.312	1.576	3.120	-60.000
.624	-3.820		

Kuznetsov, 1909

(7+1) f.t.=24° tr.t.= -1.5° (4+1)II f.t.=65°

Kremers, 1858 and 1860

%	d	%	d
	19.5°		
5	1.042	35	1.396
10	.088	40	.475
15	.138	45	.564
20	.196	50	.662
25	.258	55	.777
30	.319	60	.907

t	d		d(relative)	
			39.65%	50.74%
19.5	1.2356	1.4880	1.00000	1.00000
0	1.2414	1.5012	1.00974	1.01000
40	.2250	.4731	0.98967	0.98976
60	.2127	.4576	.97944	.97998
80	.1996	.4441	.96911	.97029
100	.1850	.4234	.95872	.96078

Kremers, 1859

d	%	t	n_D
	19.5°		
0.9983	0	17	1.3332
1.1863	-	16	.3701
.4104	-	16	.4145
.7540	-	17	.4845
2.0025	65.57	17	.5334

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.0047	1.0	1.1193
.10	.0108	2.0	.2382
.20	.0229	4.0	.4733
.50	.0590		

Herz and Martin, 1923

t	d	
	c=12.020	c=39.006
20	1.0844	1.2747
40	.0774	.2642
60	.0676	.2523
80	.0560	.2304

Jones , 1904 ; Jones and Getman, 1904

M	d	M	d
	0°		
0.078	1.014676	1.248	1.280700
.156	.032612	2.184	.484500
.312	.069476	3.120	.693268
.624	.140328		

Guillaume, 1946

%	d
20°	
11.26	1.1026
20.40	.2035
31.20	.3368

Jones, 1904 ; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.078	1.32901	1.248	1.38455
.156	.33272	2.184	.42743
.312	.34021	3.120	.46998
.624	.35553		

Herz and Martin, 1923

t	η	
	c= 12.020	c= 39.006
20	951	888
40	647	635
60	481	493
80	374	398

Jones, 1904 and Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.078	108.62	1.248	74.16
0.156	103.52	2.184	56.96
0.312	96.80	3.120	36.67
0.624	87.50		

Guillaume, 1946

%	n 5780 Å	(α)magn.10^6 5780 Å
	20°	
11.26	1.3550	4.542
20.4	1.3760	5.000
31.20	1.4036	-

In radians, gauss, centim.

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	99.5	1.0	75.5
.10	94.8	2.0	66.1
.20	89.6	4.0	49.1
.50	82.2		

Water + Calcium thiocyanate ($CaC_2N_2S_2$)

Dixon and Taylor, 1910

mol%	d	n_D
	17°	
0	0.9988	1.33319
5.40	1.2185	1.43356

Water + Calcium hydroxide (CaO_2H_2)

Lescoeur,

t	p (sat.sol.) (4+1)	(3+1)	(2+1)
10	4.8	3.35	-
15	6.5	4.70	-
20	8.7	5.20	2.9
25	10.5	-	4.5
30	13.1	9.30	5.6
35	22.0	-	-
40	22.0	14.30	-

t	p dissoc. (4+1)	(3+1)	(2+1)
10	1.1	-	-
15	-	1.0	-
20	1.9	-	-
25	-	-	-
30	4.4	3.6	2.7
35	7.3	-	-
40	9.6	6.1	-

Lunge, 1883

%	d	%	d
		15°	
1.020	1.0069	19.40	1.1247
2.250	.0140	20.55	.1335
3.480	.0121	21.70	.1425
4.800	.0285	22.94	.1516
6.070	.0358	24.01	.1608
7.340	.0434	25.49	.1702
8.460	.0509	26.58	.1798
9.600	.0587	27.86	.1896
10.780	.0665	29.11	.1994
11.970	.0744	30.34	.2095
13.150	.0825	31.55	.2198
14.440	.0907	32.83	.2301
15.680	.0990	34.04	.2407
16.930	.1074	35.43	.2515
18.170	.1160	36.78	.2624

Weir, 1955

p	Dv/Vo* at 21°			
	100	56	31	26 mol%
1143	0.0037	0.0106	0.0149	0.0173
2000	0.0000	0.0000	0.0000	0.0000
2470	-	-0.0070	-0.0115	-
3000	-0.0048	.0554	.1097	-0.1080
4000	.0073	.0678	.1206	.1366
5000	.0113	.0723	.1245	.1410
6000	.0150	.0762	.1281	.1440
7000	.0189	.0799	.1314	.1483
8000	.0217	.0830	.1350	.1517
9000	.0266	.0869	.1385	.1547
10000	.0296	.0896	.1411	.1577
	20	12	8	2.2 mol%
1143	0.0174	0.0285	0.0218	0.0227
2000	0.0000	0.0000	0.0000	0.0000
2470	-0.0111	-0.0144	-0.0175	-0.0155
3000	.1323	.1248	.0891	.0459
4000	.1390	.1328	.1018	.0633
5000	.1433	.1394	.1120	.0786
6000	.1467	.1451	.1207	.0917
7000	.1502	.1500	.1287	.1035
8000	.1528	.1564	.1362	.1139
9000	.1559	.1708	.1418	.1778

*Dv is reckoned from 1143 and from 2000 to 1000
Vo is the volume at 1 atm.

Water + Calcium nitrite (CaO_4N_2)

Oswald, 1911

%	f.t.	%	f.t.
16.7	-4	43.0	+18.5
25.5	-9.3	51.8	42
28.4	-11.5	55.2	54
29.5	-12.5	58.4	64
32.0	-14.5	60.3	70
35.0	-17.5 E	61.5	73
36.2	-9.5	71.2	91
38.3	0		

(4+1) (1+1) tr.t.: 53.5% 44°

%	d
16°	
42.3 (sat.sol.)	1.4203

Bureau, 1935 and 1937

%	f.t.	E	tr.t. I	tr.t. II
10.00	-3.75	-20.10	-	-
19.85	-9.05	-19.95	-	-
25.60	-13.00	-19.95	-	-
30.00	-16.80	-19.90	-	-
34.20	E	-20.00	-	-
35.20	-15.2(4+1)	-20.00	-	-
40.40	+2.8	"	-	-
44.85	14.0	"	-	-
50.70	28.0	"	-	-
55.05	34.6	"	34.6	-
57.15	58.0	"	34.8	-
60.2	79.5	"	34.5	-
62.5	90.8	"	34.5	-
65.70	115	-	34.0	-
71.0	- 0 aq.	-	34.0	120
73.0	-	-	34.5	128
80.4	-	-	34.4	129
86.4	-	-	34.5	126

%	t	d	%	t	d
33.01	0	1.362	57.78	64.7	1.541
45.80	18.0	.426	60.20	79.5	.549
56.25	56.8	.520	64.10	99.5	.608

Water + Calcium chlorite ($CaCl_2O_4$)

Levi and Bisi, 1956

%	f.t.	%	f.t.
11.14	0	12.25	30.5
11.55	10	12.57	40
11.92	20	12.92	50

Q diss 25° -570 cal

Water + Calcium chlorate ($CaCl_2O_6$)

Egorov, 1931

%	f.t.	%	f.t.
4.67	-1.0	51.2	-28
9.3	-2.2	-	-26.8 E
13.9	-4.2	55.0	-26.3 (4+1)
18.6	-6.3	58.0	-16.2
23.2	-9.1	60.0	-12.5
27.9	-12.9	-	-7.8 tr.t.
32.5	-17.1	63.0	-5.0 (2+1)
37.2	-24.0	66.2	+19.5
42.0	-32.9	76.2	73.5
43.0	-34.2	-	76 tr.t.
44.0	-37.0	78.0	93
45.0	-39.6	80.1	127
-	-41 E	82.3	156.5
46.0	-37.5 (6+1)	85.0	203
48.0	-30.2	92	290

Cavallaro, 1941

m	f.t.	m	f.t.
0.01032	-0.05261	0.13514	-0.64381
.02815	.13939	.19152	-0.9191
.05808	.28017	.30040	-1.4831
.07678	.36785	.44337	-2.3046
.10128	.48331	.73760	-4.1220

Rubien, 1911

N	d	n_D
	18°	
0	0.99862	1.33327
0.4855	1.03542	.33979
0.971	.0717	.34609
1.942	.1430	.35827
3.884	.2811	.38121

Heydweiller, 1912

N	d	κ	N	d	κ
			18°		
0.0971	1.00609	78.0	0.971	1.0717	589.3
.194	.01345	146.0	1.942	.1430	982.1
.485	.03542	331.5	3.884	.2811	1343.0

Heydweiller, 1913

N	n_D
	18°
0	1.33327
0.5	.33997
1.0	.34647
2.0	.35900
4.0	.38255

Water + Calcium iodate (CaI_2O_6)

Mylius and Funk, 1897

f.t.	%	f.t.	%
(6+1)		(1+1)	
0	0.1	21	0.37
10	.17	35	.48
18	.25	40	.52
30	.42	45	.54
40	.61	50	.59
50	.89	60	.65
54	1.04	80	.79
60	1.36	100	.94

Water + Calcium nitrate (CaN_2O_6)

Heterogeneous equilibria.

Wüllner, 1860

t	p 0%	p 16.7%	p 28.6%
25.00	23.55	22.05	20.25
27.62	26.45	24.65	22.60
29.80	31.14	29.14	26.74
31.80	35.00	32.66	30.35
33.60	38.62	36.27	33.50
36.00	44.20	41.11	38.20
37.43	48.73	45.64	42.63
39.75	54.16	50.82	47.39
41.00	57.91	54.18	50.64
43.20	64.20	59.55	55.95
45.70	73.98	69.28	64.89
47.50	81.14	76.14	71.14
49.40	89.21	89.88	78.35
51.20	97.62	91.66	85.70
53.40	108.75	102.20	95.75
55.65	120.89	113.34	105.89
57.45	131.73	123.33	115.34
59.40	144.69	135.95	127.21
62.80	169.25	159.12	149.38
64.22	180.37	169.45	159.22
65.10	187.79	176.53	165.96
67.60	209.89	196.79	184.08
69.40	227.12	212.71	199.07

Tammann, 1885

%	p	%	p
	100°		
6.27	746.9	32.11	647.7
11.20	735.3	37.23	614.4
15.70	719.8	41.42	583.8
18.47	710.1	42.02	578.4
21.90	697.4	44.78	556.6
26.67	676.6	62.70	356.6
30.74	655.8		

Ewing, 1927

t	p	t	p	t	p
		(4+1)+L+V			
0	2.7	39	19.5	38	8.6
5	3.9	40	19.7	37	7.75
10	5.2	41	19.7	36	7.0
15	6.9	42	19.3	35	6.4
20	9.4	42.5	19.0	34	5.75
25	12.0	42.5	15.5	32	4.7
30	14.9	42	13.8	30	3.9
35	17.7	41	12.25	25	2.4
36	18.2	40	10.9	20	1.5
37	18.9	39	9.5	15	1.15
38	19.3	-	-	-	-
	(3+1)+L+V				
35	13.3	49	20.6		
40	16.2	50	20.5		
42	17.8	50.5	20.2		
43	18.3	51	19.4		
44	18.8	51	16.8		
45	19.3	50.5	15.4		
46	19.8	50	14.4		
47	20.2	49	13.3		
48	20.5				
	(2+1)+L+V				
20	4.25	50	15.3		
30	7.1	51	15.6		
40	11.1	51.5	15.8		
45	13.2	-	-		
48	14.4	-	-		
49	15	-	-		
	0 aq.+L+V				
30	4.2	52	16.1		
35	5.9	53	17.2		
40	8.0	54	18.1		
45	10.8	55	19.0		
48	12.8	57	21.2		
50	14.6	60	24.9		
51	15.3	-	-		

t	p 20.2%	p 30.6%	p 49.7%	p 59.1%	p 66.6%
20	16.4	15.3	11.3	8.2	5.4
30	29.9	27.6	20.6	15.1	10.4
40	52.6	48.0	36.1	26.6	18.7
50	86.8	80.2	61.3	45.3	31.6
60	140.6	130.5	99.7	75.0	52.7

t	p 72.2%	p 74.2%	p 74.7%	p 76.3%	p 76.9%	p 77.4%
20	3.7	3.2	3.05	2.5	2.35	2.3
30	6.9	6.0	5.8	4.9	4.6	4.5
40	12.7	10.7	10.4	9.2	8.6	8.4
50	21.9	18.9	18.0	16.2	15.1	14.8
60	36.8	31.9	30.0	27.5	25.5	25.0

Adams and Merz, 1929

t	p
(4+1)	sat.sol.
15	7.16
20	9.73
25	12.04
30	14.88
40	19.68

Ebert, 1930

%	p 0°	10°	20°	30°
0	4.58	9.21	17.54	31.83
10	.45	9.05	17.25	31.15
20	.35	8.75	16.60	30.00
30	.15	8.25	15.45	28.15
40	3.70	7.30	13.60	24.70
50	-	5.60	11.30	20.35
60	-	-	-	15.40
satd.	2.80	5.60	9.60	15.20

Blackman and Pearce, 1936

m	p	m	p
	25°		
0.0	23.752	2.5	20.042
.1	23.659	3.0	19.107
.2	23.566	3.5	18.118
.4	23.373	4.0	17.098
.6	23.160	5.0	15.008
.8	22.915	6.0	13.062
1.0	22.638	7.0	11.260
1.5	21.868	8.0	9.603
2.0	21.002	8.3601	9.041

Keevil, 1942

mol%	t	P
43.1	329.4	18.87
65.5	451.3	65.50
76.1	500±10	35±15

Legrand, 1835

%	b.t.	%	b.t.
13.04	100	58.79	119
20.19	102	59.68	120
25.59	103	61.55	122
29.87	104	63.27	124
33.51	105	64.85	126
36.63	106	66.32	128
39.28	107	67.68	130
41.79	108	68.97	132
44.01	109	70.16	134
46.04	110	71.26	136
47.89	111	72.32	138
49.59	112	73.31	140
51.17	113	74.27	142
52.64	114	75.16	144
54.02	115	76.04	146
55.31	116	76.91	148
56.52	117	77.83	150
57.65	118	78.37	151

Gerlach, 1886

%	b.t.	%	b.t.	%	b.t.
0	100	61.08	118	74.17	136
9.09	101	62.05	119	74.74	137
16.67	102	62.96	120	75.31	138
23.08	103	63.77	121	75.87	139
28.57	104	64.59	122	76.41	140
33.33	105	65.40	123	76.98	141
37.50	106	66.16	124	77.45	142
41.18	107	66.89	125	77.97	143
44.44	108	67.57	126	78.47	144
47.09	109	68.30	127	78.95	145
49.49	110	69.00	128	79.40	146
51.58	111	69.70	129	79.90	147
53.38	112	70.38	130	80.35	148
55.05	113	71.01	131	80.78	149
56.52	114	71.67	132	81.22	150
57.90	115	72.34	133	81.63	151
59.01	116	72.97	134	82.00	151.97
60.08	117	73.57	135		

Kremers, 1856

Sat. sol. b.t. = 152°

Rüdorff, 1861

%	f.t.	%	f.t.
1.39	-0.4	10.01	-3.05
4.78	-1.4	10.60	-3.3
6.53	-1.9	12.98	-4.15

Guthrie, 1876

%	f.t.	%	f.t.
5	-1.1	20	- 6.5
10	-2.3	25	- 9.3
15	-4.2	30	-12.9

Jones, 1904 and Jones and Getman, 1904

M	f.t.	M	f.t.
0.042	-0.200	1.660	- 8.680
0.104	-0.470	2.075	-11.600
0.208	-0.910	2.905	-19.320
0.415	-1.820	3.320	-24.320
1.038	-5.070		

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.0125	-0.072	0.5	-2.2860
0.025	-0.1303	0.75	-3.484
0.05	-0.2405	1.0	-4.766
0.125	-0.5752	1.5	-7.616
0.25	-1.1424		

Jones and Stine, 1908

M	%	f.t.	M	%	f.t.
0.25	3.984	-1.189	1.05	15.33	-5.123
0.45	7.002	-2.094	1.25	17.89	-6.345
0.65	9.898	-3.041	1.45	20.35	-7.547
0.85	12.640	-4.057	1.6555	22.76	-8.859

Bassett and Taylor, 1912

%	f.t.	%	f.t.
1.40	- 0.4	68.68	42.4
4.78	- 1.4	68.74	42.5
6.53	- 1.9	- (4+1)	42.7
10.00	- 3.05	71.70	42.45
12.98	- 4.15	70.37	40.0 E
33.13	-15.70	71.45	45.0
38.70	-21.70	73.79	50.0
- E	-28.70	74.73	51.0
43.37	-26.70	- (3+1)	51.1
47.31	-10.00	77.49(2+1)	49.0
50.50	0.00	78.05	51.0
51.97	+ 5.00	78.16 0 aq.	55.0
53.55	10.00	78.20	80.0
54.94	15.00	78.37	90.0
56.39	20.00	78.43	100.0
57.98	25.00	78.57	125.0
60.41	30.00	78.80	147.5
62.88	35.00	79.00	151.0
66.21	40.00		

Taylor and Henderson, 1915

%	f.t.	%	f.t.
(4+1)I		(4+1)II	
50.17	0	61.57	30.0
56.88	22.2	63.66	34.0
57.90	25.0	67.93	39.0
60.16	30.0	69.50	39.6
62.88	35.0	75.34	39.0
64.34	38.0		
66.21	40.0		
69.50	42.7		
71.70	42.4		
77.30	25.0		

Ewing, Krey and al., 1927

%	f.t.	%	f.t.
		(3+1)	
12.5	-4.7	75.25	51.1
22.9	-9.0	76.0	51.1
33.2	-16.1	76.3	50.8
(4+1)		76.7	50
		E	50.6(3+1)+(2+1)
46.2	-15.3(I)	E	49.8(3+1)+0aq.
52.0	+7.4		
59.9	30.2	(2+1)	
62.2	35.0		
66.2	40.5	76.0	48.1
68.3	42.4	76.7	49.8
69.5	42.7	E	51.6 (2+1)+0aq.
-	39.7(II)		
69.8	42.7(I)		
70.7	42.4		
E	42.6(4+1)+(3+1)		
71.3	42.5		
71.8	42.2		
73.4	40.9		
E	39.6(4+1)+(2+1)		
76.3	36.4		
76.7	35.6		
E	32.7(4+1)+0aq.		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.827	1.6	0.917
0.2	0.819	1.8	0.934
0.3	0.818	2.0	0.953
0.4	0.821	2.5	1.001
0.5	0.825	3.0	1.051
0.6	0.831	3.5	1.103
0.7	0.837	4.0	1.157
0.8	0.843	4.5	1.210
0.9	0.850	5.0	1.263
1.0	0.859	6.0	1.361
1.2	0.879	7.0	1.452
1.4	1.898	8.0	1.535

Properties of phases.

Gerlach, 1869

%	d	%	d
	17.5°		
0	0.999	32.8	1.270
8.2	1.058	41.0	.353
16.4	.123	49.2	.443
24.6	.193		

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987	40	1.3828
10	1.0849	50	.5129
20	.1710	60	.6639
30	.1707		

Favre and Valson, 1874

m	d	m	d
	24.65°		
0	0.997	2.0	1.205
0.5	1.059	2.5	.246
1.0	.112	3.0	.286
1.5	.160	3.5	.323

Kohlrausch, 1879

%	d
18°	
6.25	1.0487
12.50	.1016
25	.2198
37.5	.3546
50	.5102

Rönnberg, 1880

t	d	t	d
0%		10%	
3.5	1.25180	3.5	1.07150
8.3	.24966	9.3	.07029
14.0	.24642	13.8	.06917
18.8	.24387	17.8	.06816
25.4	.23958	25.5	.06586
20%		30%	
3.5	1.13627	3.5	1.19620
8.8	.13455	8.5	.19416
13.8	.13268	13.8	.19171
18.3	.13077	18.3	.18942
25.8	.12742	25.5	.18539

Gerlach, 1887

%	d	%	d	%	d
		17.5°			
0	0.998	21	1.171	41	1.377
1	1.007	22	.181	42	.389
2	.014	23	.190	43	.400
3	.022	24	.200	44	.412
4	.029	25	.210	45	.423
5	.037	26	.220	46	.435
6	.045	27	.230	47	.446
7	.052	28	.239	48	.458
8	.060	29	.249	49	.469
9	.067	30	.259	50	.481
10	.075	31	.269	51	.493
11	.084	32	.280	52	.505
12	.092	33	.291	53	.518
13	.101	34	.302	54	.530
14	.110	35	.312	55	.542
15	.118	36	.323	56	.554
16	.127	37	.334	57	.566
17	.136	38	.345	58	.579
18	.144	39	.355	59	.591
19	.152	40	.366	60	.603
20	.161				

Jones, 1904; Jones and Getman, 1904

M	d	M	d
	0°		
0.042	1.006652	1.660	1.192372
.104	.012800	2.075	.236120
.208	.026220	2.905	.321880
.415	.050348	3.320	.362720
1.038	.121012		

Jones and Pearce, 1907

M	d	M	d
	20°		
0.0125	1.001846	0.5	1.06011
0.025	.003166	0.75	.08874
0.05	.00604	1.00	.11751
0.125	.01523	1.5	.17375
0.25	.03074		

Jones and Stine, 1908

M	%	d	M	%	d
		0°			
0.25	3.984	1.0310660	1.05	15.33	1.124240
.45	7.002	.054992	.25	17.89	.147288
.65	9.898	.078772	.45	20.35	.169360
.85	12.640	.100672	.6555	22.76	.193920

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	1.016	1.05972
0.1024	1.00502	2.030	.11830
0.2046	.01129	4.053	.23070
0.5101	.02963	5.504	.30758

Clausen, 1912

t	d			
	0.5154N	1.024N	2.034N	4.075N
6	1.03241	1.06342	1.12337	1.23790
18	.02995	.06018	.12010	.23172
30	.02638	.05587	.11801	.22344

Manchot, Jahrstorfer and Zepter, 1924

%	d
25°	
22.398	1.1503
44.796	1.2927

Pearce and Blackman, 1935

M	d	M	d
	25°		
0.0	0.997071	2.5	1.270014
.1	1.009670	3.0	.314619
.2	.022093	3.5	.356078
.4	.046507	4.0	.394473
.6	.070371	5.0	.462485
.8	.093683	6.0	.519536
1.0	.116449	7.0	.566558
1.5	.170956	8.0	.604500
2.0	.222150	8.3601 (satd.)	.616059

Pesce, 1935

N	d	N	d
	25°		
1.719	1.09813	6.072	1.33322
2.239	.12982	7.812	.42162
3.674	.20657	10.132	.53438

Scott and Budger, 1936

%	d	%	d
	35°		
0.0000	0.995059	32.628	1.27637
0.5282	0.997925	39.639	.35108
1.2099	1.002917	47.010	.43632
4.1147	.024410	52.39	.50280
4.8072	.029660	53.25	.51388
7.337	.049070	53.40	.51600
13.997	.102740	53.44	.51640
18.070	.137620	53.72	.52000
19.903	.153900	54.80	.53391
21.184	.165670	65.18	.67720
23.548	.187000	69.47	.74180
25.626	.206630	72.82	.79410
26.632	.216290	75.23	.83510
27.450	.224140	75.51	.83920
28.540	.235100		

Guillaume, 1946

%	d	%	d
	20°		
9.60	1.0755	25.5	1.2185
15.15	.1202	34.2	.3034
22.80	.1919	47.4	.4552

Ewing and Mikovsky, 1950

%	d 25°	30°	40°	50°	60°
0.000	0.9970	0.9956	0.9921	0.9880	0.9832
17.255	1.1364	1.1334	1.1280	1.1218	1.1154
24.476	.2024	.1992	.1928	.1862	.1792
34.052	.2984	.2949	.2881	.2806	.2728
50.358	.4850	.4808	.4728	.4645	.4561
53.108	.5206	.5161	.5076	.4990	.4905
56.201	.5608	.5565	.5481	.5392	.5304
60.263	.6158	.6114	.6030	.5938	.5857
62.765	.6505	.6459	.6374	.6285	.6195
64.320	.6726	.6683	.6596	.6505	.6420
66.122	.6981	.6934	.6847	.6759	.6671
68.124	.7299	.7259	.7168	.7080	.6991
70.469	.7654	.7607	.7520	.7428	.7351
72.503	.7989	.7941	.7854	.7770	.7677
73.396	.8132	.8096	.8010	.7921	.7831
75.145	.8431	.8384	.8302	.8207	.8124
76.815	.8704	.8653	.8578	.8491	.8402

Wagner, 1883

%	η (water°=100) 15°	25°	35°	45°
40.13	242.60	217.10	156.50	128.10
30.10	144.10	112.70	90.72	75.15
17.55	93.76	94.59	60.04	49.91

Bak, 1939

wt%	mol%	η 60°	40°	20°
8.8	1.05	519	683	1046
15.6	1.95	628	832	1304
22.7	3.12	764	1082	1633
42.7	7.56	1510	2297	4078
53.6	11.24	3571	6083	10950
61.0	14.60	7460	13741	32635
64.8	16.82	15714	37882	143160
69.5	20.00	32266	81900	344500
70.2	20.55	35060	152133	432400
72.5	22.55	48300	200200	841400
73.9	23.71	60600	214000	893000
75.5	25.26	70300	275000	1535000
76.9	26.90	146250	545000	2300000

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.042	1.32672	1.660	1.36329
.104	.32830	2.075	.37170
.208	.33070	2.905	.38707
.415	.33576	3.320	.39412
1.038	.35018	4.150	.40983

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34541
0.1	.33454	2.0	.35677
.2	.33578	4.0	.37789
.5	.33947	5.5	.39257

Faucon, 1909

%	n_D	%	n_D
	15°		
0	1.3334	8.35	1.3438
0.90	.3349	12.02	.3474
4.35	.3390	15.41	.3492

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.016	1.34561
0.1024	.33457	2.030	.35713
.2046	.33584	4.053	.37848
.5101	.33959	5.504	.39260

Muchin and Farle, 1913

M	n_D	M	n_D
	18°		
0	1.3310	1.0	1.3558
0.2	.3398	1.5	.3662
0.5	.3444	2.0	.3765

Guillaume, 1946

%	n 5780 Å	(α)* magn. 10^6
	20°	
9.6	1.3487	3.686
15.15	.3577	.516
22.8	.3712	.274
25.5	.3762	.189
34.2	.3923	2.930
47.4	.4184	.540

*in radians, gauss, centim.

de Mallemann and Guillaume, 1945

3 mol % d = 1.1899 (α)magn. = $14.82 \cdot 10^{-5}$

Kohlrausch, 1879

%	κ	$\tau.10^4$
	18°	
6.25	487	219
12.5	799	218
25	1043	219
37.5	870	254
50	466	337

Mc Gregory, 1894

N	κ	N	κ
	18°		
0.0001	0.116	0.1	84.49
0.0002	0.230	0.5	337.0
0.0006	0.677	1.0	570.2
0.001	1.117	2.0	869.0
0.002	2.195	3.0	1005.0
0.006	6.327	4.0	1046.0
0.010	10.240	5.0	1005.0
0.050	45.340	6.0	900.0

Jones, 1904 and Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.042	98.90	1.660	36.38
0.104	87.46	2.075	29.66
0.208	78.80	2.905	18.76
0.415	68.60	3.320	14.15
1.038	49.75	4.150	7.79

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.0125	108.60	0.50	66.54
0.025	102.78	0.75	58.32
0.050	96.28	1.00	51.30
0.125	85.72	1.50	40.24
0.25	77.64		

Jones and Stine, 1908

M	κ	M	κ
	0°		
0.25	209.75	1.05	563.01
0.45	332.59	1.25	603.63
0.65	429.32	1.45	632.36
0.85	502.72	1.6555	646.74

Clausen, 1912

t	$\tau.10^5$	λ
	0.5154N	
6	2512	48.74
18	3364	65.27
30	4252	82.50
	1.024N	
6	4253	41.50
18	5675	55.43
30	7176	70.08
	2.034N	
6	6484	31.88
18	8643	42.50
30	1092	53.69
	4.075N	
6	7738	18.98
18	1036	25.42
30	1319	32.37

Bak, 1939

%	κ 20°	κ 40°	κ 60°
65	85	170	330
69	50	115	215
70	55	120	220
75	10	55	105
77	20	70	115

Marignac, 1876

%	U	%	U
21-51°			
4.35	0.9510	26.71	0.7597
8.34	.9110	37.79	.6856
15.40	.8463	47.67	.6255

Rümelin, 1907

%	t	Q dil.
50.72	20.17	+57.3
50.66	20.20	+57.1
44.64	19.80	+ 1.98
44.58	19.76	+ 2.29
39.73	18.60	-23.68
39.68	18.55	-23.92
31.29	16.33	-34.80
31.23	16.30	-34.90
23.29	16.08	-27.30
23.22	16.08	-26.80
15.41	15.85	-15.68
15.32	16.13	-15.64
8.35	16.42	- 5.10
8.30	16.42	- 5.09
4.32	17.22	- 0.76
4.30	17.10	- 0.74

Water + Calcium perchlorate ($CaCl_2O_8$)

Lilich and Dzhurinskii, 1956

m	f.t.	m	f.t.
7.10	0	8.46	30
7.29	5	8.76	35
7.51	10	8.88	40
7.67	15	9.29	45
7.91	20	9.50	50
8.18	25		

Lilich and Mogilev, 1956

m	pH
25°	
0.232	5.82
0.326	5.70
0.731	5.49
1.523	4.90
5.380	3.53

Water + Monocalcium phosphate ($CaH_4O_8P_2$)

Adams and Merz, 1929

t	p
(1+1) sat.sol.	
10	8.95
15	12.67
20	16.52
25	22.89
30	29.89
40	52.37
50	87.50

Water + Calcium dithionate (CaO_6S_2)

Tammann, 1885

t	p
100°	
11.01	747.4
22.67	725.5
35.42	667.9
39.65	640.4

De Baat, 1926

%	f.t.
13.80	0
17.63	12
20.25	20
23.29	30

Water + Calcium sulfate ($CaSO_4$)

Lescoeur, 1890

t	p
(1+2)	
100	17.5
114	28
124	44.5
148	167
161.5	300

Water + Calcium thiosulfate (CaO_3S_2)

Silberman and Ivanov, 1946

f.t.	%	f.t.	%
6	28.04	25.5	34.88
9	28.84	29.9	36.77
11.75	30.08	34.5	38.48
17.75	32.28		

Water + Calcium sulfamate ($CaH_4N_2O_6S_2$)

King and Hooper, 1941

%	f.t.	%	f.t.
36.18	0.0	65.55	68.0
38.43	10.0	67.21	71.0
41.93	20.0	67.32	72.5
45.70	30.0	67.51	75.0
50.03	40.0	68.27	80.0
54.57	50.0	68.77	83.4
50.00	60.0	69.54	88.7
62.99	65.0	70.68	95.0

Schmelzle and Westfall, 1944

N	d	η (water=1)
	25°	
0.0529	1.0014	1.0140
.0826	.0051	.0275
.1618	.0169	.0448
.2209	.0148	.0502
.3432	.0248	.0826
.4657	.0336	.0991
.4718	.0372	.1378
.6014	.0454	.1580
.7029	.0529	.1711
.7258	.0594	.2310
1.0130	.0740	.4060
1.4611	.1207	.6674
2.0876	.1690	.9616
2.6746	.2188	2.5946
3.4820	.2771	

Water + Calcium formate ($CaC_2H_2O_4$)

Krasnicki, 1887

%	f.t.	%	f.t.
13.99	0	14.95	50
14.21	10	15.08	60
14.43	20	15.20	70
14.62	30	15.29	80
14.80	40		

Lumsden, 1902

%	f.t.	%	f.t.
13.90	0	14.89	60
14.86	10	15.05	70
14.23	20	15.21	80
14.40	30	15.37	90
14.57	40	15.54	100
14.74	50		

Stanley, 1904

%	f.t.
13.79	0
14.76	51.7
15.46	100

Ashton, Houston and Taylor, 1933

%	f.t.	%	f.t.
14.03	0	14.84	60
14.14	10	14.96	70
14.26	20	15.25	80
14.33	30	15.38	90
14.53	40	15.47	100
14.68	50		

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.02280	44.61
1	.04442	37.12
1.5	.06610	30.33
2	.08720	25.22

Water + Calcium acetate ($CaC_4H_6O_4$)

Krasnicki, 1887

%	f.t.	%	f.t.
23.55	80	25.47	30
24.51	70	25.79	20
25.02	60	26.46	10
25.24	50	27.60	0
25.34	40	(1+1)	

Lumsden, 1902

%	f.t.	%	f.t.
22.87	100 (1+1)	24.71	50
23.20	95	24.81	45
23.70	90	24.93	40
24.72	85	25.09	35
24.26	84 (2+1)	25.27	30
25.09	80	25.48	25
24.93	75	25.78	20
24.80	70	26.10	15
24.68	65	26.46	10
24.64	60	26.82	5
24.64	55	27.22	0

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987	20	1.1037
5	1.0247	25	.1307
10	.0516	30	.1580
15	.0775		

Hager, 1876

%	d	%	d
	17.5°		
0	0.9987		
1	1.0038	16	1.0829
2	.0090	17	.0881
3	.0142	18	.0933
4	.0194	19	.0985
5	.0247	20	.1037
6	.0300	21	.1091
7	.0354	22	.1145
8	.0408	23	.1199
9	.0462	24	.1253
10	.0516	25	.1307
11	.0568	26	.1360
12	.0620	27	.1415
13	.0672	28	.1469
14	.0725	29	.1524
15	.0778	30	.1579

Heydweiller, 1912

N	d
	18°
0.207	1.00810
0.519	.02193
1.039	.04405
2.089	.08644
3.282	.13187

de Garcia, 1920

N	d	N	d
	20°		
0.062	1.002245	0.5	1.021625
.125	.005050	1	.042570
.250	.010715	2	.082495

Guillaume, 1946

%	t	d
9.15	21.5	1.0526
18.00	20	1.1030
25.20	20	1.1330

de Garcia, 1920

N	n_D	N	n_D
	20°		
0.062	1.3341	0.5	1.3408
0.125	.3358	1.0	.3471
0.250	.3376	2.0	.3585

Guillaume, 1946

%	n 5780 Å	$(\alpha)^*_{magn.}10^6$
	20°	
9.15	1.3484	3.861
18.0	.3647	.753
25.2	.3739	.662

*in radians, gauss, centim.

Mc Gregory, 1894

N	ϰ	N	ϰ
	18°		
0.0001	0.0770	0.05	27.61
.0002	.1506	.10	49.58
.0006	.4420	.50	169.80
.0010	.7280	1.00	252.00
.0020	1.4290	2.00	316.40
.0060	4.0610	3.00	288.30
.0500	6.5510		

Heydweiller, 1912

N	ϰ
18°	
0.207	96.8
0.519	186.7
1.039	272.1
2.089	314.2
3.282	264.1

Water + Calcium propionate ($CaC_6H_{10}O_4$)

Tammann, 1885

%	p
100°	
21.48	723.7
26.44	714.9

Krasnicki, 1887

%	f.t.	%	f.t.
29.24	0	27.22	50
28.68	10	27.23	60
28.16	20	27.48	70
27.72	30	28.00	80
27.39	40		
(1+1)			

Lumsden, 1902

%	f.t.	%	f.t.	
29.97	0	27.64	55	
29.43	5	27.67	60	
29.05	10	27.75	65	
28.75	15	27.93	70	
28.49	20	28.16	75	(2+1)
28.26	25	28.49	80	
28.09	30	28.98	85	
27.93	35	29.65	90	
27.77	40	30.67	95	
27.72	45	32.63	100	
27.67	50			

Heydweiller, 1921

N	d	λ
	18°	
0.2	1.00914	43.90
0.5	.02224	32.99
1.0	.04435	23.29
2.0	.08400	12.55
3.0	.12170	6.70

Water + Calcium butyrate ($CaC_8H_{14}O_4$)

Hecht, 1882

%	f.t.	%	f.t.
16.25	0	13.19	55
15.94	5	13.08	60
15.62	10	13.04	65
15.27	15	13.04	70
15.03	20	13.04	75
14.61	25	13.04	80
14.29	30	13.07	85
13.99	35	13.12	90
13.76	40	13.26	95
13.52	45	13.65	100
13.33	50		
(1+1)			

Chancel and Parmentier, 1887

%	f.t.	%	f.t.	
16.61	0	13.18	55	
16.28	5	13.05	60	
15.76	10	13.06	65	
15.45	15	13.08	70	
15.06	20	13.12	75	
14.65	25	13.15	80	(1+1)
14.27	30	13.27	85	
13.92	35	13.42	90	
13.65	40	13.64	95	
13.43	45	13.89	100	
13.31	50			

Deszathy, 1893

%	f.t.	%	f.t.
16.88	0	13.64	50
16.00	10	13.37	60
15.22	20	13.30	70
14.58	30	13.30	80
14.04	40	(1+1)	

Lumsden, 1902

%	f.t.	%	f.t.
16.88	0	13.34	55
16.49	5	13.16	60
16.08	10	13.04	65
15.72	15	12.99	70
15.39	20	12.96	75
15.07	25	13.00	80
14.71	30	13.12	85
14.31	35	13.23	90
14.09	40	13.42	95
13.79	45	13.68	100
13.57	50		
	(1+1)		

Water + Calcium isobutyrate ($CaC_8H_{14}O_4$)

Chancel and Parmentier, 1887

%	f.t.	%	f.t.
16.90	0	20.66	55
17.15	5	21.06	60
17.41	10	21.43	65
17.70	15	21.72	70
18.00	20	21.92	75
18.33	25	21.99	80
18.68	30	21.89	85
19.04	35	21.57	90
19.42	40	20.98	95
19.82	45	20.07	100
20.24	50		
(1+1)			

Sedlitzky, 1887

%	f.t.	%	f.t.
16.87	0	20.57	50
17.46	10	21.51	60
18.13	20	22.51	70
18.78	30	23.56	80
19.69	40		

Lumsden, 1902

%	f.t.	%	f.t.
16.74	0	21.63	55
17.02	5	22.12	60
17.42	10	22.30	62
17.83	15	22.03	65
18.30	20	21.72	70
18.77	25	21.46	75
19.22	30	21.26	80
19.71	35	21.07	85
20.18	40	20.94	90
20.66	45	20.81	95
21.13	50	20.69	100
(1+1)		(5+1)	

de Garcia, 1920

N	d	n
	22.5°	
0.062	1.001696	1.3339
.125	.004421	.3352
.250	.009989	.3373
.500	.020799	.3420
1.000	.041360	.3507

Next Page

Water + Calcium Isovalerate ($CaC_{10}H_{18}O_4$)

Lumsden, 1902

%	f.t.	%	f.t.
14.20	100	18.26	45.5
14.22	95	18.23	45
14.27	90	18.03	40
14.34	85	18.89	35
14.44	80	17.81	30
14.60	75	17.81	25
14.82	70	18.89	20
15.14	65	18.13	15
15.53	60	18.50	10
15.97	55	19.17	5
16.63	50	20.66	0
(1+1)		(3+1)	

Sedlitzky, 1887

%	f.t.	%	f.t.
15.54	0	17.32	50
16.22	10	17.22	60
16.72	20	16.98	70
17.07	30	16.58	80
17.27	40		

Water + Calcium capronate ($CaC_{12}H_{22}O_4$)

Kulisch, 1893

%	f.t.	%	f.t.
10.97	0	15.77	50
13.15	10	15.04	60
14.66	20	13.74	70
15.59	30	11.80	80
15.96	40	9.09	90
(3+1)			

Water + Calcium methylethylacetate ($CaC_{10}H_{18}O_4$)

Sedlitzky, 1887

%	f.t.	%	f.t.
22.35	0	25.71	50
24.08	10	24.87	60
25.23	20	23.22	70
25.89	30	21.56	80
26.04	40		
(3+1)			

Water + Calcium methylpropylacetate ($CaC_{12}H_{22}O_4$)

Stiassny, 1891

%	f.t.	%	f.t.
14.23	0	12.23	50
13.64	10	12.12	60
13.25	20	12.11	70
12.74	30	12.20	80
12.44	40		
(3+1)			

Water + Calcium diethylacetate ($CaC_{12}H_{22}O_4$)

Keppich, 1888

%	f.t.	%	f.t.
23.26	0	18.13	40
21.76	10	17.25	50
20.40	20	16.57	60
19.18	30	16.07	70

Water + Calcium lactate ($CaC_6H_{10}O_6$)

Tammann, 1885

%	p
	100°
18.01	740.3
32.43	723.5

Water + Calcium acid malate ($CaC_8H_{10}O_{10}$)

Iwig and Hecht, 1886

%	f.t.	%	f.t.
1.27	15	24.37	57
1.36	17	11.60	68
7.85	45	6.92	78

Water + Strontium chloride ($SrCl_2$)

Heterogeneous equilibria

Tammann, 1885

t	p 0%	p 9%	p 18.41%	p 24.38%
27.29	27.2	-	24.2	22.8
34.44	40.9	-	37.4	35.3
40.92	58.1	-	51.9	49.5
44.92	71.6	-	65.8	62.3
48.20	84.6	-	77.4	73.1
51.53	99.8	96.5	91.3	86.5
53.97	112.4	109.0	102.9	97.8
56.52	127.0	123.4	116.4	110.9
61.53	160.4	155.4	146.6	139.0
65.11	188.6	183.7	173.2	164.4
67.77	212.1	206.5	194.5	185.8
71.06	244.7	228.6	224.8	213.5
74.67	285.3	277.9	262.5	250.2
78.66	336.7	327.5	308.6	293.6
81.82	382.5	372.2	352.2	334.8
82.88	415.0	403.6	381.7	363.1
88.06	488.4	475.2	449.5	428.8
90.07	527.3	513.1	485.7	462.8
93.46	599.2	583.7	552.5	526.2
95.77	652.1	632.9	600.7	571.9
100.01	760.1	749.1	702.7	669.0

100°

%	p	%	p
7.43	743.6	30.43	619.7
10.72	732.9	33.60	590.4
17.31	706.6	37.90	545.6
22.22	681.0	41.16	509.3
27.63	642.8	46.04	458.0

Hepburn, 1932

25°

%	m	p	%	m	p
3.07	0.200	23.56	26.27	2.246	20.16
4.56	.301	23.55	31.29	2.872	18.66
5.97	.400	23.35	33.97	3.245	17.70
8.69	.600	23.14	35.05	3.403	17.27
13.62	.995	22.55	35.80	3.517	16.90
19.17	1.495	21.65			

Pareau, 1877

(6+1)

t	p dissoc.	t	p dissoc.
16.9	4.3	40.6	24.3
17.2	3.2	44.0	30.0
20	5.1	44.7	31.9
29.8	10.9	44.9	32.1
29.9	11.0	50.3	46.1
36.2	17.7	55.4	63.0
40.2	23.1	55.5	63.8

Frowein, 1887

(6+1)

t	p dissoc.	t	p dissoc.
14.75	3.294	30.01	10.868
20.34	5.144	34.66	15.332
25.66	7.822	39.45	21.571

Lescoeur, 1890

t	p dissoc. (6+1)	p dissoc. (2+1)	t	p dissoc. (6+1)	p dissoc. (2+1)
5	1.7	-	30	11.0	-
10	2.4	-	40	20.1	5.6
15	3.9	-	80	192.0	69.0
20	5.6	1.8	100	409.0	235.0

Robinson, 1940

Isopiestic solutions

m_1	m_2	m_1	m_2
0.1195	0.1627	0.1574	0.2194
.5501	0.8290	0.5601	0.8434
.9452	1.5460	1.0280	1.7140
1.3170	2.3080	1.4720	2.6490
2.0430	4.0210	2.0680	4.0830
0.1935	0.2749	0.3942	0.5757
0.7120	1.1070	0.8320	1.3280
1.0940	1.7430	1.2590	2.1860
1.6820	3.1280	1.8370	3.5010
2.1570	4.3140	2.2580	4.5830

m_1 = molality of strontium chloride

m_2 = molality of potassium chloride

Stokes, 1948

25°

m	osmotic coefficient	m	osmotic coefficient
0.1	0.850	1.2	1.061
0.2	0.854	1.4	1.116
0.3	0.864	1.6	1.173
0.4	0.880	1.8	1.232
0.5	0.899	2.0	1.292
0.6	0.918	2.5	1.454
0.7	0.937	3.0	1.631
0.8	0.959	3.5	1.802
0.9	0.983	4.0	1.966
1.0	1.009		

Legrand, 1835

%	b.t.	%	b.t.
14.31	101	40.79	110
20.13	102	42.56	111
24.30	103	44.32	112
27.49	104	46.03	113
30.26	105	47.70	114
32.80	106	49.39	115
35.06	107	50.98	116
37.11	108	52.58	117
38.98	109	54.01 (satd.)	117.9

Gerlach, 1886

%	b.t.	%	b.t.
0	100	39.83	109
9.91	101	41.66	110
17.01	102	43.34	111
22.30	103	44.93	112
26.58	104	46.52	113
30.17	105	48.22	114
33.15	106	49.88	115
35.65	107	51.43	116
37.81	108	52.89	117

Baroni, 1893

%	b.t.
1.90	100.148
4.34	100.352
7.42	100.652
11.11	101.100
14.70	101.618

Kremers, 1854

%	f.t.	%	f.t.
31.15	-0.5	41.83	47
31.25	+2	44.64	57
32.57	8	47.85	67
34.72	20	49.50	84
37.17	29	50.25	93
39.22	37	52.91	106

Mulder, 1864

%	f.t.	%	f.t.
30.65	0	44.32	57.5
31.64	4.5	46.69	65 tr.t.
32.20	8.5	47.31	70.5
33.42	13.5	47.70	78.5
35.48	22.25	48.19	82.5
37.70	31	49.11	91
38.91	37.75	49.80	95
42.00	47.25	50.29	99
43.91	53	53.78	118.8 b.t.
44.29	55.3		

Rüdorff, 1872

%	f.t.	%	f.t.
5.40	-1.75	18.15	-9.00
9.90	-3.75	19.80	-10.55
13.72	-3.85	22.28	-13.00
16.97	-8.10		

de Coppet, 1872

%	f.t.
9.09	-3.05
13.04	-5.45
16.67	-7.95

Etard, 1894

%	f.t.	%	f.t.
26.5	-17	46.5	75
28.6	-11	47.1	80
29.3	-5	47.5	92
30.8	-1	49.6	98
31.3	+2	50.7	104
31.7	7	50.7	105
33.7	18	52.0	118
34.7	21.5	52.5	132
37.8	35	54.7	144
39.8	44.5	55.7	153
42.8	54	60.5	175
43.8	55	64.1	215
44.7	59	65.4	222
46.4	64	67.3	250
46.1	70		

Jones and Chambers, 1900

M	f.t.	M	f.t.
0.05	-0.258	0.30	-1.471
.10	-0.488	.40	-1.978
.135	-0.652	.50	-2.544
.20	-0.973	.75	-4.071

Jones, 1904 and Jones and Getman, 1902 and 1904

M	f.t.
1.00	- 6.000
1.50	-10.725
2.00	-16.422

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.01	-0.05273	0.10	-0.48904
0.02937	-0.1550	0.25	-1.1957
0.03987	-0.2026	0.50	-2.5339
0.5017	-0.2476	0.75	-4.0989
0.7077	-0.3472	1.00	-5.9211

Jones and Stine, 1908

M	f.t.	M	f.t.
0.15	-0.753	0.65	-3.521
0.25	-1.249	0.85	-4.863
0.35	-1.774	1.05	-6.354
0.45	-2.333	1.23541	-7.861

Menzies, 1936

m	f.t.	m	f.t.
3.43	22.9(6+1)	8.06	134.4
4.14	40.75	8.39	147.6(1+1)
4.88	54.15	8.65	156.3
5.244	58.9	9.03	169.3
5.45	60.9	9.21	175.5
5.52	62.95	9.28	175.9
5.77	75.65	9.28	177.1
6.194	93.15(2+1)	9.49	183.1
6.99	116.0	9.93	192.9
7.86	131.7		

Benrath, 1941

%	f.t.	%	f.t.
53.9	124	67.1	233
54.1	131	69.7	265
55.0	137	73.9	293
55.5	141	76.2	312
56.9	155	78.0	318
58.3	166	79.0	340
60.2	184	80.0	382
62.3	200	81.0	412

%	tr.t.
66.6	230 2-1
78.5	320 1-0

Properties of phases.

Density

Kremers, 1858 and 1859

%	d	%	d
19.5°			
0	0.9983	23.413	1.2380
8.934	1.0805	29.098	.3092
16.750	1.1612	34.076	.3793

t	d			
19.5°	1.1028	1.2039	1.2992	1.3800
0	1.1078	1.2114	1.3086	-
40	.0948	.1943	.2882	1.3680
60	.0848	.1834	.2766	.3556
80	.0732	.1714	.2641	.3425
100	.0601	.1582	.2508	.3288

Gerlach, 1859

%	d	%	d
15°			
0	0.99913	18	1.17586
1	1.00819	19	.18685
2	.01724	20	.19785
3	.02631	21	.20967
4	.03536	22	.22148
5	.04442	23	.23331
6	.05362	24	.24513
7	.06342	25	.25695
8	.07291	26	.26974
9	.08242	27	.28251
10	.09192	28	.29528
11	.10211	29	.30805
12	.11230	30	.32083
13	.12249	31	.33458
14	.13268	32	.34833
15	.14287	33	.36207
16	.15387		
17	.16486		

Fouqué, 1867

%	d°	t	d^t
0	1.0154	12	1.0139
1.99	.0159	9	.0157
5.23	.0435	12	.0420
17.84	.1559	12	.1530
100	.1565	10	.1538

Favre and Valson, 1874

m	d	m	d
	24.7°		
0	0.997	2.0	1.247
0.5	1.067	2.5	.301
1.0	.130	3.0	.352
1.5	.190	3.5	.401

Kohlrausch, 1879

%	d
18°	
5	1.0443
10	.0932
15	.1456
20	.2023
22	.2259

Volkmann, 1882

%	d
15-16°	
0	0.9990
6.28	1.0567
12.79	.1204
22.56	.2282
29.25	.3114

Schumann, 1887

%	d
15°	
0	0.9991
1.24	1.0100
6.43	1.0580
17.70	1.1726
28.26	1.2858

Jahn, 1891

c	d
20°	
0	0.9982
10.621	1.0875
23.484	1.1919

Charpy, 1893

%	d	%	d
	0°		
0	0.9999	23.2300	1.2513
6.7243	1.0636	27.7170	1.3084
12.9997	1.1283	31.8193	1.3607
18.2629	1.1913		

Sentis, 1897

mol %	t	d
0	13.5	-
2	10.6	1.1469
4	10.7	1.2828

Grabowsky, 1904

%	d 10°	d 18°	d 30°
0	0.999731	0.99863	0.99567
8.49	1.0808	1.0784	1.0748
16.02	1.1597	1.1565	1.1516
22.75	1.2356	1.2315	1.2253

Chéneveau, 1907

%	d	%	d
	13.5°		
0	0.99935	8.89	1.0805
3.12	1.0261	11.56	1.1087
6.08	1.0544	14.11	1.1348
7.50	1.0672		

Jones and Pearce, 1907

M	d	M	d
	20°		
0.01	1.0012284	0.10	1.013205
0.02937	.0038396	0.25	.034433
0.03987	.0053832	0.50	.068379
0.5017	.007028	0.75	.101760
0.7077	.00956	1.00	.135423

Schneider, 1910

N	d	N	d
	18°		
0	0.99862	0.4925	1.03281
0.0985	1.00561	0.9850	1.06623
0.197	1.01241	2.0	1.1348

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	1.005	1.06773
0.1020	1.00597	1.995	.13411
.2028	.01292	3.977	.26282
.5013	.03379	5.489	.36015

Herz, 1914 and 1917

N	d(1914)	d(1917)
25°		
0	0.9971	0.9971
1.125	1.0722	1.0718
2.250	.1456	.1453
3.375	.2176	.2179
4.500	.2879	.2880

Herz and Hiebenthal

N	d	
	25°	70°
0.5	1.0319	1.0130
1.0	.0658	.0461
2.0	.1310	.1104
4.0	.2604	.2366

Pesce, 1934

N	d	N	d
	25°		
1.375	1.06978	2.209	1.14424
.157	.07524	3.698	.23937
.240	.08067	4.355	.28057
.855	.12124	5.496	.35126

Okazaki, 1935

%	d	%	d
	28°		
6.61	1.0567	25.83	1.267
13.20	.1228	30.90	.332
13.64	.1274	34.10	.378
20.91	.2080		

Guillaume, 1946

%	d
20°	
8.60	1.0807
11.56	.0983
25.50	.2672
33.40	.3736

Schneider, 1910

N	η(water=1)	N	η(water=1)
	18°		
0.0985	1.0155	0.985	1.1380
0.1970	.0289	2.0	1.3082
0.4925	.0669		

Herz, 1914 and 1917

N	η(water=1)	η (1917)
25°		
1.125	1.190	1065
2.250	.400	1253
3.375	.690	1513
4.500	2.077	1859

Herz and Hiebenthal, 1929

N	η	
	25°	70°
0.5	958.7	446.1
1.0	1026.2	483.2
2.0	1183.8	570.5
4.0	1774.0	813.8

Volkmann, 1882

%	σ
15-16°	
0	73.2
6.28	75.0
12.79	76.5
22.56	79.6
29.25	82.5

Sentis, 1897

mol %	t	σ
0	13.5	74.0
2	10.6	77.6
4	10.7	82.4

Grabowsky, 1904

%	σ 10°	σ 30°
0	74.00	71.02
8.49	75.92	72.91
16.02	77.73	74.75
22.75	78.84	75.96

Kremers, 1856 and 1857

%	t	n_D
0	16	1.3333
19.93	14	.3715
31.97	16	.4001
sat.sol.	b.t.= 114°	

Börner, 1869

t	$n_{H\alpha}$	t	$n_{H\beta}$	t	$n_{H\gamma}$
		9.9938%			
44.10	1.344716	45.75	1.350881	45.0	1.354418
39.35	.345488	40.70	.351649	39.95	.355256
35.25	.346098	35.15	.352543	35.6	.355934
29.30	.346923	29.65	.353329	28.7	.356914
24.60	.347532	24.70	.353954	24.8	.357538
		19.9876%			
45.05	1.359531	45.9	1.366205	45.2	1.369956
40.0	.360332	40.65	.367091	39.95	.370840
35.7	.361043	34.95	.368050	36.0	.371581
30.4	.361878	30.8	.368666	30.45	.372428
24.9	.362659	24.8	.369568	24.7	.373240
		29.951%			
45.25	1.373204	46.0	1.380294	45.2	1.384256
42.2	.373681	41.6	.381066	40.6	.385148
35.2	.374879	35.3	.382171	34.75	.386127
30.85	.375601	31.6	.382732	29.0	.387036
25.75	.376340	25.6	.383626	25.5	.387631

Jahn, 1891

c	n Hα	n D	n Hβ
0	1.3315	1.3332	1.3375
10.621	.3493	.3512	.3558
23.484	.3690	.3712	.3761

Borgesius, 1895

%	n_D
	20°
0	1.333000
1.23	.333221
4.73	.333890
10.46	.335101
16.39	.336538

Chéneveau, 1907

%	n_D	%	n_D
	13.5°		
0	1.33365	8.89	1.3496
3.12	.3388	11.56	.3549
6.08	.3443	14.11	.3601
7.50	.3470		

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34711
0.1	.33471	2.0	.36036
.2	.33613	4.0	.38549
.5	.34033	5.5	.40352

Jahn, 1891

c	(α)magn.
	20°
0	1
10.621	1.0828
23.484	1.1116

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.005	1.34718
0.1020	.33474	1.995	.36027
.2028	.33617	3.977	.38518
.5013	.34035	5.489	.40339

Guillaume, 1946

%	$(\alpha)^*_{magn.}10^6$ 5780 Å	n
8.6	4.046	1.3497
11.56	4.058	.3536
25.5	4.127	.3795

* in radians, gauss, centim.

Kohlrausch, 1879

%	ϰ	τ
	18°	
5	470	0.0215
10	881	.0209
15	1223	-
20	1486	-
22	1573	-

Mc Gregory, 1894

N	ϰ	N	ϰ
	18°		
0.0001	0.120	0.1	90.28
.0002	0.236	0.5	376.1
.0006	0.690	1.0	679
.0010	1.137	1.5	938
.0020	2.236	2.0	1150
.0060	6.442	2.5	1329
.0100	10.400	3.0	1473
.0500	46.820	3.5	1592

Jones, 1904, Jones and Getman, 1902 and 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.50	84.26	1.50	63.01
1.00	74.57	2.00	55.26

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.00091	128.57	0.07127	98.06
0.00182	127.99	0.100	95.98
0.01000	115.64	0.250	88.07
0.02937	106.10	0.500	82.18
0.03987	103.26	0.751	74.86
0.05018	101.04	1.000	71.23

Jones and Stine, 1908

M	ϰ	M	ϰ
	0°		
0.15	153.77	0.65	550.17
0.25	241.85	0.85	678.24
0.35	325.90	1.05	797.96
0.45	404.79	1.23541	886.07

Pesce, 1934

N	ϰ	N	ϰ
	25°		
1.075	1164	2.209	1266
1.157	1174	3.698	1376
1.240	1185	4.355	1421
1.855	1235	5.496	1486

Okazaki, 1935

%	Verdet's constant 3441 Å
20°	
6.61	0.04782
13.20	.05164
13.64	.05172
20.91	.05679
25.83	.06011
30.90	.06370
34.10	.06599

Okazaki, 1942

%	Verdet's constant (D)
25°	
10.29	0.01456
14.36	.01524
18.43	.01592
21.46	.01649
24.40	.01702
28.90	.01790
32.08	.01847

Mc Clung and Mc Intosh, 1902

d	X-ray absorption (in%)
room temperature	
1.250	94.9
.130	90.2
.032	83.8
.015	78.0
.007	75.6
.000	65.3

Heat constants .

Stearn and Smith, 1920

m	g.sol	moles water added	Q dil. (by mole aq.)
3.2	10.0	35.0	-5.4
"	"	15.0	6.4
2.9	9.5	37.2	3.49
"	"	30.0	3.75
"	"	30.0	3.80
"	"	27.7	3.95
"	"	20.0	4.01
"	"	19.5	4.38
"	"	10.0	4.85
2.4	10.0	35.0	0.96
"	"	25.0	1.29
2.0	"	35.0	1.01
"	"	25.0	1.23
1.55	"	40.0	0.59
"	"	30.0	0.71
"	"	30.0	0.70
"	"	20.0	0.86
1.2	"	35.0	0.54
"	"	25.0	0.65
0.8	"	40.0	0.38
"	"	30.0	0.44
"	"	20.0	0.50
0.425	"	40.0	0.28
"	"	30.0	0.34
"	"	20.0	0.37
0.2	"	40.0	0.20
"	"	30.0	0.23
"	"	20.0	0.24

Lehtonen, 1922

m	Q diss.(cal/g)
0°	
0.0625	+80.490
.1250	80.168
.2500	79.711
.5000	78.871
1.0000	77.278
2.0000	75.051
4.0000	71.107

Water + Strontium bromide ($SrBr_2$)

Heterogeneous equilibria .

Tammann, 1885

t	p 0%	15.67%	26.63%	33.37%	45.86%
46.01	75.7	72.6	68.5	62.7	50.8
49.79	91.6	86.7	82.6	75.6	61.3
54.08	113.0	107.8	102.2	94.5	77.1
57.11	130.6	126.2	118.9	109.4	89.5
60.61	153.7	147.4	138.9	128.4	105.4
63.41	174.7	169.1	159.1	143.8	120.3
66.23	198.2	191.6	180.4	165.8	136.4
69.83	232.1	223.2	210.7	194.2	160.0
73.04	266.4	255.9	241.3	222.7	184.0
75.21	292.9	281.1	265.3	243.1	202.7
78.14	329.5	317.3	299.2	275.8	228.4
80.56	363.6	348.8	330.1	304.4	253.0
84.87	431.5	413.8	391.1	361.8	301.5
87.87	484.9	466.6	441.3	408.2	340.3
91.18	550.0	527.7	499.4	461.3	385.4
94.23	616.2	592.2	560.4	518.5	433.0
97.95	706.0	679.0	643.3	594.4	497.7
100.52	774.2	744.3	702.4	653.2	548.5

%	p	%	p	%	p
100°					
8.13	747.7	28.66	684.3	45.57	546.1
13.82	736.0	33.29	656.6	48.92	502.9
21.17	714.9	38.11	621.3	53.07	444.1
24.37	703.0	41.23	595.6	53.08	443.8

Lescoeur, 1890

t	p dissoc. (6+1)	sat.sol.
20	1.7	9.1
40	5.4	28.8
100	190	190

Robinson, 1942

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.0896	0.1256	1.079	1.940
0.1166	0.1648	1:095	1.993
0.1953	0.2834	1.119	2.050
0.2413	0.3519	1.161	2.147
0.3056	0.4548	1.178	2.192
0.3323	0.4941	1.306	2.496
0.5257	0.8255	1.508	3.007
0.5829	0.9330	1.610	3.431
0.6524	1.055	1.659	3.431
0.7331	1.210	1.758	3.699
0.8362	1.427	1.942	4.249
0.9517	1.672	2.123	4.795
0.9518	1.668		

m_1 = m of $SrBr_2$ m_2 = m of KCl

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.859	0.9	1.042
0.2	0.871=	1.0	1.074
0.3	0.888	1.2	1.142
0.4	0.908	1.4	1.210
0.5	0.932	1.6	1.284
0.6	0.957	1.8	1.360
0.7	0.983	2.0	1.440
0.8	1.011		

Johnston, 1906

%	b.t.	%	b.t.
14.43	100.857	62.79	113.32
20.48	101.645	67.45	119.94
31.95	103.152	70.45	125.03
38.46	104.770	72.42	127.81
45.71	107.526	74.49	131.12

Etard, 1894

%	f.t.	%	f.t.
43.1	-11	51.8	20
46.7	-1	68.5	93
47.0	+1	68.7	97
48.2	+7	69.8	107
51.7	18		

Jones and Chambers, 1900 and
Jones and al., 1904, 1905 and 1910

M	f.t.	M	f.t.
0.053	-0.262	0.517	-2.741
.103	.503	.621	3.447
.155	.773	.626	3.502
.207	-1.035	1.565	12.520
.259	.308	1.878	16.500
.310	.592	2.191	-22.000
.414	-2.147		

Milikan, 1914

%	f.t.	%	f.t.
6.9	-1.7	29.8	-13.5
13.9	-4.2	34.8	-19.0
23.2	-8.4	41.7E ice+(6+1)	-28.0

Benrath, 1941

%	f.t.	%	f.t.
69.4	104	82.7	303
71.9	154	84.4	319
73.0	176	89.0	331
75.1	212	90.0	335
76.2	224	91.0	338
78.2	249	92.0	342
78.7	262	93.3	366
81.6	291	94.0	383

tr.t. (1+1) - 0 aq. 92.4% 345°

Properties of phases .

Kremers, 1856, 1858 and 1859

%	d	%	d
	19.5°		
5	1.044	30	1.330
10	1.092	35	1.408
15	1.144	40	1.490
20	1.202	45	1.587
25	1.264	50	1.691

t	d		
19.5°	1.4250	1.6001	1.2039
0	1.4361	1.6139	1.2106
40	.4220	.5849	.1943
60	.3982	.5692	.1829
80	.3831	.5529	.1696
100	.3670	.5356	.1545

Jahn, 1891

c	%	d
	20°	
0	0	0.9982
17.003	14.90	1.1416
34.994	27.12	1.2899

Jones and al., 1904, 1905 and 1910

M	d	M	d
	0°		
0.053	1.009752	0.517	1.107480
.103	.022160	.621	.129680
.155	.031528	.626	.127052
.207	.041988	1.565	.315820
.259	.054596	1.878	.378784
.310	.064152	2.191	.441612
.414	.086584		

Heydweiller, 1909

M	d	M	d
	18°		
0.05	1.0040	1.0	1.1033
.10	.0093	2.0	.2062
.20	.0199	4.0	.4113
.50	.0516		

Thomas and Perman, 1934

%	d	%	d
	30°		
13.03	1.126	42.17	1.518
24.70	.262	50.24	.655
33.98	.387	55.60	.763

Guillaume, 1946

%	d
20°	
14.91	1.1430
20.30	.2086
35.0	.4081

Thomas and Perman, 1934

%	π	%	π
	30°		
13.03	38.8	42.17	27.9
24.70	34.2	50.24	24.8
33.98	30.9	55.60	22.6

Kremers, 1857

%	n_D
17°	
0	1.3332
32.20	.3912
48.56	.4375

Jahn, 1891

c	n H_α	n D	n H_β	(α)magn.
0	1.3315	1.3332	1.3375	1
17.003	.3552	.3573	.3622	1.3192
34.994	.3765	.3820	.3875	1.3273

Borgesius, 1895

%	t	D n_D(.10^6)
1.90	22.6	276
7.15	22.6	1127
17.55	22.4	3148
22.95	22.6	4489

Guillaume, 1946

%	(α)magn.10^6 5780 Å	n 5780 Å
	20°	
14.91	4.180	1.3585
20.3	4.244	1.3686
35.0	4.404	1.4004

(α)magn. in radians, gauss, centim.

Jones and Chambers, 1900

M	λ	M	λ
	25°		
0.518	70.66	0.155	85.8
0.259	79.63	0.103	90.7
0.207	82.85		

Johnston, 1906

N	λ	N	λ
	99.4°		
0.001	366.1	0.5	205.0
0.010	311.6	3	128.4
0.100	251.2	5	95.1
0.250	228.4	7	69.3

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	100.4	1.0	71.1
0.10	93.8	2.0	62.0
0.20	86.6	4.0	45.6
0.50	78.5		

Water + Strontium iodide (SrI_2)

Robinson, 1942

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.873	0.9	1.139
0.2	0.897	1.0	1.179
0.3	0.926	1.2	1.266
0.4	0.956	1.4	1.357
0.5	0.988	1.6	1.452
0.6	1.024	1.8	1.548
0.7	1.060	2.0	1.649
0.8	1.099		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.876	0.9	1.136
0.2	0.899	1.0	1.177
0.3	0.925	1.2	1.264
0.4	0.955	1.4	1.352
0.5	0.987	1.6	1.443
0.6	1.021	1.8	1.540
0.7	1.056	2.0	1.641
0.8	1.095		

Etard, 1894

%	f.t.	%	f.t.
60.0	-20	68.5	63
60.3	-10	70.5	77
62.2	-3	74.0	81
63.0	+7	79.2	97
63.4	11	79.4	105
63.5	18	80.8	120
64.8	38	85.6	175
66.0	52		

Jones and Getman, 1904

M	f.t.	M	f.t.
0.027	-0.140	0.216	-1.156
0.054	-0.275	0.281	-1.505
0.070	-0.359	0.327	-1.804
0.081	-0.415	0.562	-3.656
0.108	-0.558	1.125	-9.040
0.141	-0.719	1.687	-17.000
0.162	-0.844	2.812	-49.500

M	d	M	d
	0°		
0.070	1.020504	0.562	1.159796
0.141	.039092	1.125	.318728
0.281	.079098	1.687	.477628

Kremers, 1858 and 1860

%	d	%	d
	19.5°		
0	0.998	35	1.408
5	1.043	40	1.489
10	1.089	45	1.587
15	1.140	50	1.692
20	1.198	55	1.809
25	1.260	60	1.952
30	1.328	65	2.146

t	d				
19.5	1.2321	1.4489	1.6725	1.8752	2.0573
0	1.2397	1.4611	1.6884	1.8936	-
40	.2217	.4348	.6551	.8557	2.0360
60	.2095	.4199	.6376	.8364	2.0155
80	.1958	.4040	.6198	.8170	1.9949
100	.1809	.3874	.6016	.7977	1.9768

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.00594	1.0	1.1420
0.10	.01313	2.0	.2035
0.20	.02753	4.0	.5618
0.5	.07050		

Kremers, 1859

%	t	n_D
0	17	1.3332
42.37	18	1.4227
61.72	17	1.5026

Jones and Getman, 1904

M	n_D	molecular conducitvity	M	n_D	molecular conductivity
"		0°			
0.070	1.32901	117.8	1.125	1.38319	84.69
0.141	.33272	108.0	1.687	.41223	71.70
0.281	.34004	99.9	2.812	.46761	44.57
0.562	.35463	96.4			

Heydweiller, 1909

N	λ	N	λ
	18°		
0.05	99.0	1.0	75.2
0.10	93.9	2.0	66.1
0.20	88.5	4.0	47.9
0.50	81.6		

Water + Strontium oxide (SrO)

Lescoeur, 1890

t	p dissoc. (2+1)	(9+1)	sat.sol.
5	-	2.15	5.65
10	-	3.30	8.20
15	-	5.00	-
20	-	7.50	-
30	-	13.10	-
40	-	21.20	-
85	96	-	-
100	239	735	-

Scheibler, 1861

%	f.t.	%	f.t.
0.35	0	2.54	55
.41	5	3.03	60
.48	10	3.62	65
.57	15	4.35	70
.68	20	5.30	75
.82	25	6.56	80
1.00	30	9.00	85
.22	35	12.00	90
.48	40	15.15	95
.78	45	18.60	100
2.13	50	19.40	101.2

Scheibler and Sidersky, 1881 and Sidersky, 1922

%	f.t.	%	f.t.
0.35	0	2.94	59
.47	11	3.74	65
.52	13	4.03	67
.69	20	4.16	68
.70	21	4.20	69
.77	24	5.29	75
1.48	40	6.24	79
.54	41	6.91	81
.66	43	9.08	85
.69	44	10.74	88
.91	48	11.95	90
2.13	50	19.34	101.2
2.80	57		

Grube and Nussbaum, 1928

%	f.t.	%	f.t.
0.30	0	11.92	90
0.79	25	15.00	95
1.26	35	17.35	98
2.04	50	18.56	100
5.34	75	(9+1)	

Water + Strontium nitrite (SrN_2O_4)

Oswald, 1914

%	f.t.	%	f.t.
11.3	-1.3	46.5	52.5
19.6	-3.1	48.2	56
32.8	-6.8 E	49.3	60.5
33.4	-2.3(1+1)	50.7	65.5
34.5	-0.3	54.0	82.5
35.5	-7.7	56.6	92
39.3	+19	58.1	98
43.1	35		

39.3% sat sol. b.t.= 112.5/763 mm d^{19}=1.4459

Bureau, 1937

%	f.t.	E	tr.t.	
11.91	-3.25	-8.75	-	
19.9	-6.20	"	-	
26.4	-8.80	-8.80	-	
35.05	+7.0	-8.75	-	(4+1)
39.55	15.0	-8.80	15.0	
40.65	16.80	-8.80	15.05	(1+1)
55.0	94.0	-8.75	15.0	
68.8	-	-8.80	15.0	
75.0	-	-	15.0	

%	t	d
32.25	1.0	1.363
41.30	29.0	1.445
43.5	36.0	1.475
47.60	56.0	1.526
51.9	79.5	1.595
56.85	99.5	1.671

Water + Strontium chlorite ($SrCl_2O_4$)

Levi and Bisi, 1956

%	f.t.	%	f.t.
0	13.05	45	14.92
15	14.00	55	15.32
23	14.15	65	15.72
25	14.08	75	16.10
35	14.51		

Q diss. 25° -1140 cal.

Water + Strontium chlorate ($SrCl_2O_6$)

Cavallaro, 1941

m	f.t.	m	f.t.
0.01032	-0.05263	0.13514	-0.64624
.02815	.14083	.19152	-0.91412
.05808	.28343	.30040	-1.4473
.07678	.37199	.44337	-2.1725
.10128	.48715	.73760	-3.9520

Linke, 1953

%	f.t.		%	f.t.	
9.29	-1.8		55.0	-38.1	
17.92	-4.0		57.1	-28	(3+1)
29.94	-9.4		57.9	-16.7	
37.46	-14.2		59.9	-7	
45.11	-22.2		61.4	0	
49.87	-29.4		63.47	+9	
62.23	0	0aq.	65.32	65	
63.55	15		66.0	75	
63.78	25		66.07	80	
64.17	35		66.41	85	
64.55	45		67.68	95	
64.65	50		67.8	110	
64.95	55		68.7	121	

E: 54.5% -37.0° tr.t.: 63.4% +10°

Traube, 1895

%	d	%	d
	15°		
0	0.99913	10.376	1.08604
3.900	1.03056	12.791	1.10786
7.263	1.05877	18.859	1.16579

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	0.9920	1.09666
0.0996	1.00877	1.981	1.1918
0.2017	1.01894	3.948	1.3760
0.4981	1.04826		

Heydweiller, 1912

N	d	N	d
	18°		
0.0988	1.00857	0.998	1.0961
0.1975	1.01840	1.975	1.1916
0.4940	1.04050	3.951	1.3763

Grufki, 1913

N	d	N	d
	18°		
0	0.99862	1.9730	1.19157
0.5043	1.04873	4.0000	1.38097
1.0340	1.10063		

Linke, 1953

%	d	%	d
0	1.828	65	1.842
10	.829	75	.845
15	.830	80	.847
25	.831	85	.849
35	.833	95	.853
45	.835	110	.861
50	.837	121	.867
55	.838		

E: 54.5% -37.0° tr.t.: 63.4% +10°

Grufki, 1913

N	n H_α	n H_β	n H_γ
		18°	
0	1.33145	1.33737	1.34059
0.5043	.33856	.34463	.34793
1.034	.34591	.35206	.35541
4.000	.38423	.39106	-

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34730
0.1	.33475	2.0	.36047
.2	.33618	4.0	.38516
.5	.34041		

Heydweiller, 1912

N	κ	N	κ
	18°		
0.0988	78.3	0.998	572.5
.1975	146.3	1.975	933.5
.494	324.5	3.951	1206.0

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.9920	1.34719
0.0996	.33474	1.981	.36022
0.2017	.33619	3.948	.38454
0.4981	.34036		

Water + Strontium bromate ($SrBr_2O_6$)

Heydweiller, 1921

N	d	λ
	18°	
0.2	1.02887	63.9
0.5	.0717	54.8
1.0	.1423	46.42
1.5	.2112	40.05

Water + Strontium nitrate (SrN_2O_6)

Heterogeneous equilibria .

Tammann, 1885

%	p	%	p
	100°		
6.80	749.0	31.19	690.9
11.37	741.1	33.08	684.5
15.68	732.7	37.91	665.0
26.27	705.9	44.00	636.9

Gerlach, 1886

%	b.t.	%	b.t.
0	100	42.06	103.5
10.71	100.5	44.87	104
19.35	101	47.26	104.5
25.82	101.5	49.39	105
31.03	102	51.22	105.5
35.23	102.5	52.86	106
38.87	103	53.81	106.3

Buchanan, 1899

%	b.t.	%	b.t.
	760.0 mm		
50.02	106.32	39.07	103.95
49.72	106.53	36.20	103.46
46.00	105.44	32.98	102.96
43.99	104.94	29.27	102.47
41.77	104.45		
	741.2 mm		
41.86	104.36	19.79	100.71
36.50	102.84	17.50	100.50
33.40	102.34	15.10	100.30
30.00	101.83	13.82	100.20
25.76	101.32	12.37	100.10
23.74	101.12	11.14	100.01
21.83	100.91	9.59	99.91

Kremers, 1856

Sat. sol. b.t. = 108°

Berkeley and Applebey, 1911

Sat. sol. b.t. = 106.790°

Kremers, 1854

%	f.t.	%	f.t.
30.12	0	49.50	50
36.63	10	50.25	75
47.62	25	51.55	100

Mulder, 1864

%	f.t.	%	f.t.
28.32	0	47.51	34.5
30.79	4.25	48.50	60.8
37.54	13.50	49.68	68
38.80	16.50	49.01	78
44.38	25	50.10	98
46.49	29.75	50.58	107.9 b.t.
47.20	31		

tr.t. = 31.3°

Rüdorff, 1872

%	f.t.	%	f.t.
7.41	-1.5	23.08	-5.45
13.79	3.0	24.24	5.75
16.67	3.7	25.37	6.10
19.35	4.45	26.47	-6.44
21.88	-5.0		

de Coppet, 1872

%	f.t.	%	f.t.
9.09	-2.0	23.08	-5.3
16.67	-3.7	25.93	-6.25
20.00	-4.6		

Etard, 1894

%	f.t.	%	f.t.
24.5	-6	47.8	56
35.9	+14	49.1	76
39.8	20	50.4	94
46.9	32	50.2	110
47.2	53		

Jones, 1904 . Jones and Getman, 1904
and Jones and Bassett, 1905

M	f.t.	M	f.t.
0.145	-0.68	1.451	-5.85
0.290	-1.35	1.741	-7.13
0.580	-2.44	1.8137	-7.50
1.161	-4.70		

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.01	0.05717	0.25	1.0817
0.025	0.1304	0.5	2.0849
0.05	0.2402	0.75	3.0453
0.075	0.3492	1.00	3.9983
0.10	0.4587		

Berkeley and Applebey, 1911

%	f.t.	%	f.t.
0.58	28.63	39.74	47.39
14.71	37.84	47.73	47.76
26.40	45.07	61.34	48.42
29.06	46.70	68.96	48.87
30.28	46.97	78.98	49.46
32.58	47.07	88.94	50.03

Sieverts and Petzold, 1933

%	f.t.
5.0	- 0.8
7.5	- 1.4
12.7	- 2.5
21.7	- 4.55
24.7	- 5.4 E ice + (4+1) I
28.9	- 6.6 (4+1) II
31.4	- 7.6
35.1	- 8.8
37.3	- 9.8
43.9	-13.0
41.9	-13.2
42.6	-13.9
44.4	-14.5
28.2	+ 0.1 (4+1) I
40.7	+20.0
45.8	+28.0
47.0	+29.3 tr.t. (4+1) I - 0 aq.
47.2	+35.0
48.3	+60.0
59.2	+80.0
51.2	105.0
46.6	-13.7 0 aq. II
45.7	- 5.0

Benrath, 1942

%	f.t.	%	f.t.
55	179	75	370
60	225	80	425
65	265	85	475
70	312		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.816	1.2	0.754
0.2	0.796	1.4	0.754
0.3	0.785	1.6	0.754
0.4	0.778	1.8	0.755
0.5	0.773	2.0	0.758
0.6	0.769	2.5	0.768
0.7	0.765	3.0	0.783
0.8	0.762	3.5	0.800
0.9	0.760	4.0	0.818
1.0	0.757		

Properties of phases .

Kremers, 1855 - 1856

%	d	%	d	%	d
			19.5°		
0	0.998	14	1.120	28	1.266
1	1.007	15	.129	29	.278
2	.015	16	.138	30	.290
3	.023	17	.148	31	.302
4	.032	18	.158	32	.314
5	.039	19	.168	33	.328
6	.047	20	.179	34	.338
7	.057	21	.190	35	.352
8	.066	22	.200	36	.365
9	.074	23	.211	37	.379
10	.083	24	.221	38	.393
11	.093	25	.233	39	.408
12	.101	26	.244	40	.420
13	.111	27	.255		

Favre and Valson, 1874

m	d	m	d
	23.4°		
0	0.997	2.0	1.289
0.5	1.081	2.5	1.350
1.0	1.155	3.0	1.407
1.5	1.224		

Kohlrausch, 1879

%	d	%	d
	18°		
5	1.0445	20	1.2047
10	1.0939	24	1.2559
15	1.1473		

Long, 1880

%	d	%	d
	15°		
5	1.0418	20	1.1815
10	.0857	25	.2363
15	.1318	35	.3542

Kanonnikoff, 1885

%	d
22.2°	
0	0.9977
18.02	1.1583

Gerlach, 1887

%	d	%	d
	17.5°		
0	0.999	30	1.292
10	1.082	40	1.420
20	1.178		

Le Blanc and Rohland, 1896

%	d
20°	
0	0.9982
10.50	1.0885
25.51	1.2419

Jones and Pearce, 1907

M	d	M	d
	20°		
0.01	1.001525	0.25	1.04201
0.025	.004207	0.5	.08312
0.05	.008391	0.75	.12386
0.075	.012646	1.00	.16354
0.10	.016834		

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	1.014	1.34643
0.1029	1.33471	2.019	1.35861
0.2051	1.33606	4.029	1.38109
0.5080	1.34004		

Jones and Getman, 1904
Jones, 1904 and Jones and Bassett, 1905

M	d	M	d
	0°		
0.145	1.026100	1.451	1.301120
0.290	1.050300	1.741	1.379600
0.580	1.100608	1.8137	1.400900
1.161	1.231000		

Berkeley and Appleby, 1911

%	t	d	%	t	d
28.63	0.58	1.28561	47.39	39.74	1.51282
37.84	14.71	.39380	47.76	47.73	.51150
45.07	26.40	.48831	48.42	61.34	.51048
46.70	29.06	.51098	48.87	68.96	.51057
46.97	30.28	.51441	49.46	78.98	.51091
47.07	32.58	.51408	50.03	88.94	.51174

Okazaki, 1935

%	d	%	d
	28°		
7.35	1.057	27.63	1.2637
14.42	1.120	34.19	1.3439
21.49	1.191		

Guillaume, 1946

%	d	%	d
	20°		
14.75	1.1282	27.90	1.2682
24.75	1.2323	33.99	1.3419

de Lannoy, 1895

t	relative volume			
	4%	8%	13%	25%
0	1.00000	1.00000	1.00000	1.00000
10	.00108	.00170	.00288	.00380
20	.00298	.00417	.00635	.00810
30	.00600	.00750	.01040	.01273
40	.00980	.01160	.01485	.01754
50	.01415	.01619	.01986	.02298
60	.01930	.02140	.02535	.02868
70	.02497	.02750	.03130	.03468
80	.03153	.03420	.03760	.04090

Wagner, 1883

%	η (water°=1) 15°	25°	35°	45°
32.61	1.1690	0.9333	0.7675	0.6233
21.19	0.8730	.6918	.5782	.4808
10.29	0.6935	.5598	.4593	.3908

Kanonnikoff, 1885

%	n H_α	n D	n H_β
	22.2°		
0	1.33108	1.33288	1.33716
18.02	.354812	.357406	.362125

Walter, 1889

c	n_D	c	n_D
	15°		
0	1.3334	21.7	1.3639
3.18	1.3389	37.9	1.3911
8.94	1.3457		

Doumer, 1892

%	n_D	%	n_D
	15°		
9.51	1.3470	45.32	1.3965
18.02	1.3588	53.08	1.4074
27.56	1.3722	60.03	1.4173
37.21	1.3852		

Borgesius, 1895

c	D n_D (sol. - aq.)
6.574	0.000845
25.968	0.003346

Le Blanc and Rohland, 1896

%	n_D
20°	
0	1.3333
10.50	1.3471
25.51	1.3695

Jones, 1904; Jones and Getman, 1904; Jones and Bassett, 1905

M	n_D	M	n_D
	0°		
0.073	1.32711	1.161	1.35347
.145	.32949	.451	.36053
.290	.33420	.741	.36701
.580	.34021		

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34625
0.1	.33467	2.0	.35838
.2	.33599	4.0	.38075
.5	.33993		

Faucon, 1909

%	n_D	%	n_D
	15°		
0	1.3334	10.51	1.3458
1.16	.3351	14.98	.3490
5.54	.3416	19.02	.3519

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.014	1.34643
0.1029	.33471	2.019	.35861
.2051	.33606	4.029	.38109
.5080	.34004		

Guillaume, 1946

%	t	(α)* magn. 10^6 5780 Å	n
14.75	20	3.528	1.3544
24.75	18	3.218	.3693
27.90	19	3.119	.3756
33.99	19	2.933	.3854

* in radians, gauss, centim.

Long, 1880

%	κ	τ.10⁴	%	κ	τ.10⁴
		18°			
5	294	225	20	762	228
10	501	225	25	823	226
15	656	227	35	818	241

Mc Gregory, 1896

N	κ	N	κ
	18°		
0.0001	0.111	0.01	9.78
0.0002	0.220	0.05	42.61
0.0006	0.647	0.10	79.38
0.0010	1.068	0.50	306.80
0.0020	2.095	1.00	508
0.0060	6.017	3.00	856

Jones and Getman, 1904

Jones, 1904 and Jones and Bassett

M	molecular conductivity	M	molecular conductivity
	0°		
0.073	90.86	1.161	40.60
0.145	80.90	1.451	34.00
0.290	76.60	1.8137	30.00
0.580	56.70		

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.010	111.14	0.25	73.59
0.025	100.95	0.50	61.34
0.050	94.12	0.75	52.77
0.075	89.26	1.00	45.69
0.10	86.16		

Okazaki, 1935

%	Verdet's constants.10^5 3514 Å	%	Verdet's constants.10^5 3514 Å
	28°		
7.35	4183	27.63	4130
14.42	4183	34.19	4122
21.49	4164		

de Mallemann and Guillaume, 1945

d = 1.2659 20° (α)magn.10^5 = 19.26

Jager, 1891

%	relative heat conductivity coeff.
0	100
20	96.4
36	92.3
40	92.8

Water + Strontium perchlorate ($SrCl_2O_8$)

Lilich and Dzhurinskii, 1956

m	f.t.	m	f.t.
8.16	0	10.54	25
8.57	5	11.43	30
9.03	10	12.02	35
9.49	15	12.70	40
10.10	20		

Pesce, 1935

N	d	N	d
	25°		
0.892	1.09153	3.619	1.37055
1.046	.12085	4.001	.40857
1.336	.13783	7.463	.74659
2.765	.28460		

Lilich and Mogilev, 1956

m	pH	m	pH
	25°		
0	6.95	2.02	5.20
0.155	5.84	2.74	4.85
0.914	5.48	5.05	4.44
1.384	5.21		

Water + Strontium acid sulfate ($SrH_2O_8S_2$)

Wasif, 1953

M	κ	M	κ
	25°		
0.0842	192.3	0.5765	488.2
.1705	285.1	.6750	508.6
.2853	382.4	.8069	509.5
.3844	436.0	.9290	502.8
.4666	465.0		

Water + Strontium dithionate (SrO_6S_2)

Tammann, 1885

%	p	%	p
		100°	
13.21	750.6	33.36	711.1
20.67	741.5	36.75	698.9
31.51	718.8		

De Baat, 1926

%	f.t.	%	f.t.
4.51	0	10.80	20
7.37	12	14.90	30

Water + Strontium formate ($SrC_2H_2O_4$)

Stanley, 1904

%	f.t.	%	f.t.
7.71	0	17.86	67.5
8.71	11	20.86	81.5
11.70	28.6	21.61	86.0
12.80	37.4	21.30	91.7
15.13	51.4	21.08	100

Ashton, Hous and Saylor, 1933

%	f.t.	%	f.t.	
	(2+1)			
8.33	0	20.00	60	
9.43	10	23.08	70	
11.18	20	23.72	72	
13.00	30	24.20	80	0 aq.
15.07	40	24.76	90	
17.35	50	25.43	100	

Heydweiller, 1921

N	d	λ
	18°	
0.2	1.01416	60.7
0.5	.0349	49.89
1	.0689	40.28
1.5	.1024	33.56

Water + Strontium acetate ($SrC_4H_6O_4$)

Osaka and Abe, 1911

%	f.t.	%	f.t.
		(4+1)	
26.97	0.05	29.97	9.0
28.52	5.0	30.36	10.0
29.50	7.5		
		(1+2)	
30.31	8.0	27.73	40.0
30:19	9.0	27.38	44.98
30.05	10.0	27.19	50.0
29.96	11.0	26.83	60.0
29.73	13.0	26.60	70.0
29.53	15.0	26.52	80.0
28.98	20.0	26.46	85.0
28.67	25.0	26.60	90.0
28.30	30.0	26.62	93.0
27.97	35.03	26.67	97.0

tr.t. = 9.4°

Rubien, 1911

N	d
18°	
0	0.99862
0.2000	1.01260
0.5018	1.03296
1.003	1.06660
1.472	1.09702
2.936	1.18806

Heydweiller, 1912

N	d
18°	
0.200	1.01260
0.5018	1.03296
1.003	1.06640
1.472	1.09702
2.930	1.18805

de Garcia, 1920

N	d
26°	
0.062	1.001910
0.125	1.006091
0.250	1.014370
0.500	1.030660
1.000	1.063105

Guillaume, 1946

%	d
20°	
11.96	1.0856
15.98	1.1160
25.55	1.1930

Rubien, 1911

N	n_D
18°	
0	1.33327
0.2000	1.336356
0.5018	1.34069
1.003	1.34782
1.472	1.35421
2.936	1.37275

Heydweiller, 1913

N	n_D
18°	
0	1.33327
0.2	1.33636
0.5	1.34066
1.0	1.34778
2.0	1.36112
3.0	1.37355

de Garcia, 1920

N	n_D
26°	
0.062	1.3331
0.125	1.3341
0.250	1.3361
0.500	1.3396
1.000	1.3463

Guillaume, 1946

%	(α)magn.10^6 5780 Å	n 5780 Å
11.96	3.774	1.3513
15.98	3.699	1.3577
25.55	3.521	1.3733

(α)magn. in radians, gauss, centim.

Mc Gregory, 1894

N	$\varkappa$	N	$\varkappa$
18°			
0.0001	0.0818	0.05	31.21
0.0002	0.1633	0.10	56.65
0.0006	0.4863	0.50	201.10
0.001	0.8045	1.00	309.5
0.002	1.568	2.00	371.8
0.006	4.521	3.00	388.7
0.010	7.308		

Heydweiller, 1912

N	$\varkappa$
18°	
0.200	101.2
0.5018	200.7
1.003	306.9
1.472	358.6
2.930	360.8

Water + Barium chloride ($BaCl_2$)

Heterogeneous equilibria.

Tammann, 1885

t	p 0%	9.98%	23.97%	24.32%	28.77%
50.53	95.0	93.0	87.0	87.4	86.1
55.59	121.5	119.6	112.3	112.1	110.4
59.08	143.2	140.6	132.3	132.5	130.8
61.25	158.3	155.4	145.9	146.0	144.0
64.53	183.7	179.5	169.0	168.8	166.2
68.73	221.2	216.5	203.2	203.7	200.3
70.65	240.5	234.2	221.4	222.6	216.1
73.55	272.6	265.8	251.2	250.9	245.6
76.26	305.0	297.9	281.5	280.6	275.1
79.86	353.4	345.4	326.1	325.2	319.0
82.84	398.3	388.3	367.4	366.9	359.1
84.87	431.7	421.2	398.4	397.7	389.7
88.01	487.6	477.7	450.9	449.3	441.1
90.07	527.4	516.0	487.5	486.1	476.2
92.46	577.0	564.5	532.5	532.2	420.3
95.35	642.3	627.7	594.4	592.3	580.1
97.91	705.0	687.0	652.4	649.5	635.7
100.55	775.0	755.3	715.3	713.8	699.5

%	p
100°	
10.45	741.9
17.75	721.7
26.58	692.6
30.42	677.2
33.22	664.0

Biltz, 1902

%	p
25°	
0	23.76
2.63	23.49
5.775	23.40
8.05	23.23
17,11	22.58
17.15	22,58

Hepburn, 1932

m	%	p
	25°	
0.199	3.98	23.56
0.289	5.68	23.53
0.406	7.80	23.52
0.508	9.57	23.28
0.606	11.21	23.20
1.007	17.34	22.77
1.285	21.11	22.42
1.568	24.61	21.82
1.655	25.64	21.68
1.788	27.14	21.53

Pareau, 1877

t	p dissoc.	t	p dissoc.
	(2+1)		
18.1	1.8	51.9	32.5
18.4	2.6	52.8	33.8
25.7	2.9	53.1	34.1
29.0	3.9	53.1	34.2
29.2	4.0	54.0	36.9
43.9	14.1	55.5	41.8
50.9	29.2		

Frowein, 1887

t	p dissoc.	t	p dissoc.
	(2+1)		
18.25	2.970	36.45	12.745
25.68	5.461	36.85	13.114
25.90	5.548	39.30	13.478
28.85	7.125	43.45	21.117
31.65	8.945		

Lescoeur, 1890

t	p dissoc. (6+1)	(2+1)	(1+1)
5	5.4	-	-
10	7.5	-	-
30	-	5.7	-
40	-	10.5	4
60	-	60	20.5
80	-	208	50.5
100	-	623	271

Robinson, 1940

Isopiestic solutions

m_1	m_2	m_1	m_2
	25°		
0.08810	0.1209	0.1134	0.1558
.1553	.2139	.3213	.4526
.5660	.8278	.7809	1.175
1.322	2.148	1.405	2.294
.497	2.472	1.649	2.773
0.1475	0.2034	0.1529	0.2106
.4649	.6696	.5088	.7360
.9302	1.436	1.227	1.962
1.454	2.395	1.492	2.462
1.782	3.046		

m_1 = molality of Barium chloride

m_2 = " " Potassium chloride

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.843	0.8	0.904
0.2	0.837	0.9	0.918
0.3	0.843	1.0	0.934
0.4	0.853	1.2	0.967
0.5	0.864	1.4	1.002
0.6	0.877	1.6	1.034
0.7	0.890	1.8	1.065

Legrand, 1835

%	b.t.	%	b.t.
0	100.2	27.85	102.7
9.91	100.7	30.79	103.2
16.01	101.2	33.46	103.7
20.79	101.7	35.90	104.2
24.53	102.0	37.54	104.6

De Heen, 1881

%	b.t.
11.48	100.9
16.39	101.6
22.40	102.4

Gerlach, 1886

%	b.t.	%	b.t.
0	100	24.01	102.5
6.01	100.5	27.38	103
11.26	101	30.41	103.5
15.97	101.5	33.11	104
20.19	102	35.56	104.5

Buchanan, 1899

%	b.t.	%	b.t.
	non satd.		
7.54	100.00	16.37	100.72
9.14	100.11	18.30	100.92
10.54	100.21	20.06	101.12
11.85	100.31	21.74	101.32
12.95	100.41	23.24	101.52
14.28	100.51	26.49	102.03

%	b.t.	%	b.t.
	758.64 mm		
19.30	101.52	32.50	103.54
23.30	102.02	35.07	104.05
26.82	102.53	36.87	104.43
29.71	103.03	36.91	104.46
	620.56 mm		
14.59	95.62	26.47	97.14
17.67	95.82	28.64	97.54
20.03	96.12	30.98	98.04
22.21	96.42	33.79	98.65
24.22	96.74	33.99	98.80
	549.94 mm		
26.46	93.83	34.07	95.32
29.20	94.34	34.13	95.43
31.64	94.84		

Kahlenberg, 1901

%	b.t.	%	b.t.
	742.9 mm		
3.92	99.606	24.25	101.607
8.07	99.889	28.81	102.334
11.84	100.189	32.37	102.989
15.70	100.538	35.28	103.521
19.40	100.950		
	752.8 mm		
3.28	99.942	25.95	102.251
7.66	100.230	27.97	102.611
11.98	100.573	30.67	103.096
16.58	101.027	32.49	103.471
21.30	101.643	35.00	103.891

Johnston, 1906

%	b.t.	%	b.t.
2.03	100.197	29.15	102.854
3.85	100.313	33.19	103.413
10.04	100.738	37.64	104.134
16.61	101.308	41.17	104.528
24.07	102.058	44.73	104.548

Kremers, 1856

Sat.sol. b.t.= 106°

Mulder, 1864

%	f.t.	%	f.t.
23.61	0	30.55	49
24.24	4.25	33.90	78
24.70	8	37.62	104.1 b.t.
28.21	34.5		

Gérardin, 1865

%	f.t.
24.92	9
27.59	30
28.57	37
30.41	50
31.46	58

Nordenskjöld, 1869

%	f.t.	%	f.t.
23.73	0.0	34.17	77.5
25.31	12.2	36.59	95.65
29.18	38.4	37.07	102.5
32.29	62.75	37.38	105.0

Rüdorff, 1872

%	f.t.	%	f.t.
0.84	-0.2	14.80	-4.20
1.68	.4	15.38	4.50
3.28	.75	16.50	5.00
4.82	1.15	17.05	5.30
6.32	.50	18.65	5.90
7.75	.90	19.68	6.45
9.14	2.30	20.66	6.75
10.48	2.65	21.63	7.30
11.76	3.1	22.10	7.70
13.00	3.4	22.56	8.00
14.21	-3.95	23.48	-8.45

de Coppet, 1872

%	f.t.	%	f.t.
4.76	-1.02	19.35	-5.15
9.09	-2.05	20.17	-5.75
12.28	-2.85	21.88	-6.60
16.67	-4.2	23.08	-7.50

Guthrie, 1876

%	f.t.	%	f.t.
5	-0.9	21.83	-7.5 E
10	-2.2	23.98	0.0 (1+1)
15	-4.0	30	+25
20	-6.0		

Etard, 1894

%	f.t.	%	f.t.
36.0	100	38.9	160
37.3	130	40.7	180
37.5	144	43.1	215

Kistiakowsky, 1890

%	f.t.
2.43	-0.59
4.74	-1.16
9.17	-2.35
17.02	-5.10

Jones and Chambers, 1900, Jones, 1904

M	f.t.
0.0976	-0.473
.248	0.932
.2929	1.413
.4882	2.418
.5858	-2.945

Jones and Getman, 1902 and 1903, Jones, 1905

M	f.t.	M	f.t.
0.25	-1.227	0.05	+0.238
.50	2.511	.10	.485
.75	3.907	.20	.954
		.30	1.446
		.50	2.490
		.75	3.851

Findlay and Cruickshank, 1925

26.32% f.t.= 20°

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
(2+1)		(1+1)	
2.6	9.7	4.8	102.8
2.7	29.2	4.9	113.1
3.7	53.1	5.1	127.8
3.8	65.1	5.4	153.0
4.2	78.3	5.6	168.7
4.8	98.5	6.1	180.0

Properties of phases.

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0164	16	1.1575
4	.0348	18	.1804
6	.0536	20	.2041
8	.0732	22	.2288
10	.0932	24	.2544
12	.1139	26	.2809
14	.1353	26.7	.2904

Kremers, 1855

%	d
19.5°	
8.156	1.0742
15.426	.1502
21.587	.2225
26.166	.2815

Schiff, 1858-1859

%	d	%	d
	21.5°		
0	0.9979	13.64	1.1280
0.85	1.0053	14.50	.1371
1.71	.0127	15.35	.1465
2.56	.0202	16.20	.1561
3.41	.0278	17.05	.1660
4.26	.0353	17.90	.1760
5.12	.0431	18.75	.1860
5.97	.0509	19.61	.1962
6.82	.0589	20.46	.2066
7.67	.0671	21.31	.2173
8.53	.0755	22.16	.2280
9.38	.0839	23.02	.2388
10.23	.0925	23.87	.2498
11.09	.1012	24.72	.2611
11.94	.1100	25.59	.2725
12.79	.1189		

Rönnberg, 1880

t	d	t	d
	5 %		
4.2	1.04371	20.7	1.04092
10.0	1.04308	25.5	1.03962
15.5	1.04215		
	10 %		
4.0	1.08673	20.7	1.08288
10.0	1.08564	25.7	1.08125
15.5	1.08432		
	15 %		
4.0	1.12822	20.7	1.12373
10.2	1.12685	26.0	1.12180
15.5	1.12538		
	20 %		
4.2	1.16866	20.7	1.16362
10.5	1.16702	25.5	1.16174
15.5	1.16545		
	25 %		
4.2	1.20836	20.5	1.20259
10.0	1.20660	26.0	1.20016
15.5	1.20462		

Gerlach, 1859

%	d	%	d
15°			
0	0.99913	14	1.13679
1	1.00830	15	.14746
2	.01745	16	.15897
3	.02661	17	.17050
4	.03577	18	.18202
5	.04493	19	.19354
6	.05477	20	.20506
7	.06461	21	.21786
8	.07444	22	.23065
9	.08428	23	.24346
10	.09412	24	.25626
11	.10480	25	.26906
12	.11546		
13	.11613		

Fouqué, 1867

%	d 0°	d 12°
0.90	1.0078	1.0076
4.42	.0371	.0362
15.53	.1439	.1417

Kohlrausch and Grotrian, 1875

%	d
18°	
5	1.0446
10	.0940
15	.1475
20	.2051
24	.2564

Forster, 1878

%	t	d
0	24.0	0.9973
11.13	24.0	1.1061
21.01	23.0	1.2101

Volkmann, 1882

%	d	%	d
15-16°		21°	
0	0.999	0	0.9980
5.48	1.0497	6.13	1.0544
10.04	.0947	12.50	.1190
18.00	.1822	19.36	.1951
26.00	.2820	24.17	.2561

Bender, 1883

M	d
15°	
0	0.9991
1	1.0895
2	.1780
3	.2647

Blümcke, 1884

%	t	d	%	t	d
0	15.0	0.999	14.85	16.2	1.135
5.12	15.4	1.042	20.23	15.0	.202
9.92	13.0	1.089	23.80	13.7	.244

Schumann, 1887

%	d	%	d
15°			
0	0.9991	10.84	1.1031
2.11	1.0184	18.44	.1889
4.52	1.0406	20.16	.2071

Bender, 1887

N	d	$\tau.10^6$	
	15°	20-25°	15-20°
1	1.0894	1470	1291
1.5	.1334	1608	1452
2	.1780	1730	1591
2.5	.2211	1825	1704
3	.2647	1904	1807

Brückner, 1891

mol%	d	mol%	d
15°			
0	0.9992	2	1.1765
0.5	1.0447	2.5	.2212
1	.0891	3	.2654
1.5	.1334		

Jahn, 1891

c	d
20°	
0	0.9982
15.818	1.1337
34.613	1.2896

Semenov, 1893

%	t	d	%	t	d
0	15	0.9991	20.0	16.25	1.1813
2.5	15	1.0191	22.5	16.25	.2057
4.0	16.8	.0288	25.0	15.60	.2343
5.0	15.6	.0392	27.5	20.00	.2619
7.5	15.0	.0611	30.0	15.90	.2850
10.0	14.4	.0833	32.5	20.00	.3150
12.5	16.25	.1038	35.0	15.00	.3591
15.0	15.0	.1295	37.5	20.00	.3880
17.5	16.25	.1518	40.0	13.75	.4327

Charpy, 1893

%	d	%	d
0°			
0	0.9999	14.0898	1.1426
5.0532	1.0476	18.0040	.1890
9.7746	1.0957	21.6511	.2333

Grabowsky, 1904

%	d 10°	d 18°	d 30°
0	0.999731	0.99863	0.99567
10.23	1.0984	1.0955	1.0911
18.32	.1894	.1858	.1804
24.62	.2674	.2632	.2595

Chéneveau, 1907

%	d	%	d
13.5°		15°	
0	0.99935	0	0.9991
4.79	1.0430	5.35	1.0480
9.21	.0859	10.24	.0968
11.29	.1088	12.53	.1214
13.30	.1298	14.72	.1452
16.79	.1716	18.84	.1927
20.61	.2122	22.64	.2392

Schneider, 1910

N	d	N	d
18°			
0	0.99862	0.493	1.0438
0.0985	1.00774	0.990	.0884
0.1197	1.01682	1.990	.1773

Rubien, 1911

N	d	N	d
18°			
0	0.99862	1.0125	1.08949
0.1025	1.00807	2.0208	.17770
.2048	.01743	3.1364	.27554
.5072	.04457		

Clausen, 1912

N	d 6°	d 18°	d 30°
0.501	1.04659	1.04448	1.04105
1.000	.09227	.08936	.08535
1.996	.18174	.17785	.17274
3.106	.28074	.27491	.26933

Herz, 1914 and Herz, 1917

N	d	d (1917)
25°		
0	0.9971	0.9971
0.614	1.0511	1.0521
1.228	.1057	.1060
1.842	.1590	.1588
2.456	.2122	.2055

Manchot, Jahrstorfer and Zepter, 1924

c	d
	25°
12.914	1.1090
13.135	.1085
27.328	.2266
27.348	.2290

Kohner, 1928

m	d
	25°
0	0.99707
0.24989	1.04192
0.49935	.08551
1.00079	.16960
1.46700	.24404

Jones and Doll, 1929

M	d	M	d
	25°		
0.004996	0.99823	0.09997	1.01497
.005000	.99819	.24990	.04212
.010035	.99887	.24948	.04210
.009990	.99881	.4972	.08663
.024988	1.00161	.4959	.08640
.024995	.00160	.9913	.17358
.048598	.00587		
.049960	.00615		

Herz and Hiebenthal, 1929

N	d	
	25°	70°
0.5	1.0418	1.0234
1.0	.0863	.0407
2.0	.1738	.1517

Shibata and Hölemann, 1931

m	d	m	d	m	d
25°		35°		45°	
0	0.99707	0	0.99406	0	0.99024
0.4045	1.06903	0.50845	1.07982	0.40015	1.06036
.53625	.09176	.6706	.11067	0.74805	.11894
.83425	.14210	.86795	.14343	1.30505	.20858
1.1805	.19867	1.3144	.21520	1.5293	.24326
.3062	.21876	1.5245	.24794		
.528	.25352				

Gibson, 1935

%	d	%	d
	25°		
0.00	0.9971	15.00	1.1441
5.00	1.0424	19.70	.1982
9.96	1.0909	24.79	.2620

Spacu and Popper, 1935

%	d	%	d
	21°		
0	0.99802	12.121	1.11408
3.998	1.03428	16.915	.16706
7.976	1.07244	21.085	.21676

Okazaki, 1935

%	d
	28°
6.81	1.0582
12.22	.1122
13.95	.1307
20.31	.2034
25.14	.2641

Guillaume, 1946

%	d
	20°
11.0	1.1043
24.6	1.2609

Kremers, 1856

t	relative volume		
	10.47%	19.48%	26.09%
0	0.99597	0.99423	-
10	0.99764	0.99695	0.99655
19.5	1.00000	1.00000	1.00000
30	.00336	.00388	.00419
40	.00715	.00797	.00847
50	.01152	.01247	.01305
60	.01647	.01732	.01793
70	.02188	.02258	.02313
80	.02777	.02823	.02853
90	.03416	.03427	.03435
100	.04098	.04069	.04043

De Heen, 1881

t	r.v.	t	r.v.	t	r.v.
22.40%		16.39%		11.48%	
10.00	1.000000	10.00	1.000000	10.00	1.000000
15.25	.001717	13.85	.001115	11.50	.000329
20.32	.003448	19.22	.002751	16.10	.001454
27.05	.005941	27.28	.005545	21.15	.002828
36.20	.009628	35.50	.008679	27.95	.005014
44.45	.013117	43.84	.012186	36.20	.008011
52.40	.016750	51.12	.015507	45.15	.011705
58.90	.019970	60.35	.020035	55.15	.016346
67.33	.024405	70.36	.025248	63.72	.020667
75.00	.028749			73.43	.025992

Kurochkin, 1929

c*	Dv%
4	-0.96
10	2.40
20	4.80

* g(1+1) in 100cc

Schumann, 1887

%	d	%	d
15°			
0	0.9991	10.84	1.1031
2.11	1.0184	18.44	.1889
4.52	1.0406	20.16	.2071

Schmidt, 1859

%	t	π
2	16.2	46.3
5	16.4	44.6
11.5	15.5	41.2
15	14.7	39.6
18	14.6	36.7

Gibson, 1935

%	π (1-1000 bars)
0.00	39.35
5.00	37.19
9.96	35.12
15.00	33.12
19.70	31.19
24.79	29.00

Viscosity and surface tension

Grotrian, 1876

%	t	η	t	η
5.25	2.34	2474	15.80	1565
15.08	2.47	2677	15.66	1900
23.56	2.61	3120	14.93	2207

%	η	τ.10⁴
18°		
5.25	1416	247
15.08	1762	204
23.56	1979	213

Brückner, 1891

mol%	η 15°	η 20°	mol%	η 15°	η 20°
0	1143.9	1008.6	2	1423.8	1272.2
0.5	1201.8	1064.5	2.5	1520.3	1361.5
1	1268.7	1128.5	3	1638.3	1468.8
1.5	1341.3	1196.7			

Schneider, 1910

N	η(water=1)	N	η(water=1)
18°			
0.0985	1.0133	0.99	1.1207
.1197	.0282	1.99	1.2666
.493	.0611		

Herz, 1914

N	η(water=1)	N	η(water=1)
25°			
0	1	1.842	1.277
0.614	1.077	2.456	1.384
1.228	1.159		

Herz, 1917

N	η	N	η
25°			
0	895	1.842	1143
0.614	964	2.456	1239
1.228	1037		

Mazetti, 1924

N(18°)	t	η	N(18°)	t	η
0.486	13	1241	0.972	13	1294
	34.4	805		34.4	838
	45.7	679		45.6	697
	59	554		59	578
1.457	13.2	1351	1.943	13	1429
	34.4	882		34.4	930
	45.6	733		45.7	775
	59	601		59	639
2.429	13.2	1509			
	34.4	991			
	45.9	823			
	59	670			

Herz and Hiebenthal, 1929

N	η 25°	η 70°
0.5	850.5	444.6
1.0	1005.4	469.3
2.0	1135.3	552.5

Jones and Dole, 1929

M	η(water=1)	M	η(water=1)
	25° (different series)		
0.004996	1.00244	0.09997	1.02678
.005000	.00252	.09997	.02707
.010035	.00370	.09997	.02701
.009990	.00372	.24990	.06448
.024988	.00807	.24990	.06424
.024995	.00835	.24948	.06441
.048598	.01456	.4972	.12878
.048598	.01470	.4972	.12891
.049960	.01483	.4959	.12876
.049960	.01486	.9913	.28059
.09996	.02716	.9913	.28039
.09996	.02710		

Volkmann, 1882

%	σ	%	σ
21°		15-16°	
0	72.4	0	73.2
6.13	73.8	5.48	74.7
12.50	75.4	10.04	75.5
19.36	76.5	18.00	76.9
24.17	77.7	26.00	78.8

Ochsé, 1890

t \ c	a^2 0	5	10	15	20
4	80.7	-	-	-	-
5	-	64.3	-	-	53.0
6	-	-	58.3	53.8	-
15	75.9	62.4	57.5	-	52.0
17	-	-	-	54.7	-
25	73.0	60.7	55.2	53.7	51.2
35	-	58.3	53.2	52.9	50.6
40	66.8	-	-	-	-
45	-	56.2	52.2	52.1	49.6
55	62.7	55.1	51.0	51.3	48.8

Grabowsky, 1904

%	σ 10°	σ 30°
0	74.00	71.02
10.23	75.62	72.54
18.32	76.94	74.09
24.62	78.34	75.48

Optical and electrical properties .

Kremers, 1856 and 1857

%	n_D		
	16°		
0	1.3333		
24.30	1.3738	sat.sol.	b.t. = 106°

Börner, 1869

t	$n_{H\alpha}$	t	$n_{H\beta}$	t	$n_{H\gamma}$
		9.9935%			
44.9	1.341360	45.35	1.347610	44.8	1.351119
40.4	.342152	39.95	.348524	39.2	.351977
35.2	.342007	35.25	.349222	34.65	.352652
30.4	.343554	30.6	.349870	31.0	.353228
27.95	.343910	28.0	.350225	28.1	.353585
		19.9872%			
46.7	1.353925	45.95	1.360713	45.2	1.364424
41.2	.354871	40.1	.361726	40.2	.365222
35.3	.355834	35.65	.362365	35.4	.366037
30.7	.356547	30.8	.363235	30.8	.366675
		29.9811%			
47.15	1.365842	48.55	1.372579	47.2	1.376631
42.0	.366675	43.0	.373566	42.75	.377441
37.8	.367454	37.8	.374448	38.3	.378250
32.6	.368268	32.4	.375347	33.0	.378988
28.65	.368905	28.8	.375910	28.95	.379761

Forster, 1878

%	t	n_D
0	24.0	1.3326
11.13	24.0	.34970
21.01	23.0	.36772

Jahn, 1891

c	$n_{H\alpha}$	n_D	$n_{H\beta}$
		20°	
0	1.3315	1.3332	1.3375
15.818	.3536	.3556	.3602
34.613	.3779	.3801	.3852

Borgesius, 1895

%	t	Dn_D (sol.-aq.)
1.37	21.6	0.000199
5.25	21.0	.000825
17.97	20.8	.003278
20.17	16.4	.003822

Chéneveau, 1907

%	n_D	%	n_D
13.5°		15°	
0	1.33365	0	1.3334
4.79	.3405	5.35	.3412
9.21	.3474	10.24	.3493
11.29	.3510	12.53	.3532
13.30	.3542	14.72	.3571
16.79	.3614	18.84	.3651
20.61	.3682	22.64	.3726
18.7°	19.74	1.1988	

%	n_C	n_D	n_{Tl}	n_F	$n_{G'}$
			18.7°		
19.54	1.36362	1.36571	1.36789	1.37038	1.37424

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.0125	1.34834
0.1025	.33486	2.0208	.36257
.2048	.33640	3.1364	.37804
.5072	.34092		

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34831
0.1	.33483	2.0	.36260
.2	.33634	3.0	.37643
.5	.34088		

Kohner, 1928

m	n_D
25°	
0	1.33253
0.24989	.34003
0.49935	.34717
1.00079	.36082
1.46700	.37259

Shibata and Hölemann, 1931

m	$n_{He\ y}$	m	$n_{He\ y}$	m	$n_{He\ y}$
25°		35°		45°	
0	1.33270	0	1.33149	0	1.33000
0.4045	.34463	0.50845	.34565	0.40015	.34161
.53625	.34833	.6706	.35064	0.74805	.35114
.83425	.35647	.86795	.35592	1.30505	.36545
1.1805	.36552	1.3144	.36735	1.5293	.37089
.3062	.36871	1.5245	.37248		
.528	.37410				

Spacu and Popper, 1935

%	$n_{He\ y}$	%	$n_{He\ y}$
	21°		
0	1.332568	12.121	1.351854
3.998	.338648	16.915	.360208
7.976	.344951	21.085	.368180

Jahn, 1891

c	$(\alpha)_{magn.}$
20°	
0	1
15.818	0.94625
34.613	0.93740

Guillaume, 1946

%	t	n 5780Å	$(\alpha)^{*}_{magn.}10^{6}$
11.0	17	1.3506	3.968
24.6	20	.3760	.934

*in radians, gauss, centim.

Barbier and Roux, 1890

c	t	dispersive power 4524 Å-6452 Å
10.4	24.1	0.370
15.6	24.8	.382
20.8	24.8	.395

Dewar and Fleming, 1897

t	ε	t	ε
	23%		
-205.0	2.62	-111.0	35.20
-198.2	2.52	-101.7	69.00
-173.7	2.64	-84.8	164.20
-145.0	8.20		

Kohlrausch and Grotrian, 1875

%	κ 0°	κ 18°
5	249	387
10	476	729
15	695	1045
20	885	1317
24	1933	1525

Grotrian, 1876

%	κ	τ.10⁴
	18°	
5.25	405	209
15.08	1050	196
23.56	1504	190

Kohlrausch, 1879

%	κ	τ.10⁴
	18°	
5	387	215
10	730	686
15	1045	983
20	1324	1245
24	1525	1435

Bender, 1887

N(15°)	κ
18°	
0.5	383
1	695
1.5	959

Mc Kie, 1918

N	κ	N	κ
	18°		
0.1915	162.8	2.029	1211
.3044	249.0	.317	1323
.5060	400.5	.775	1474
.6796	500.0		

Mazzetti, 1924

N(18°)	t	κ	N(18°)	t	κ
0.486	13.2	350.6	0.972	13.2	640.0
	34.4	517.9		34.4	929.5
	45.5	605.6		45.4	1083.3
	59.0	713.4		58.6	1262.6
1.457	13.2	889.5	1.943	13.1	1101.1
	34.1	1279.4		34.4	1581.9
	45.5	1494.9		45.6	1838.4
	58	1728.1		59.1	2137.2
2.429	13.5	1295.7			
	34.5	1838.3			
	45.6	2123.4			
	59.1	2466.2			

Jones and Chambers, 1900; Jones, 1904

M	λ	M	λ
25°		0°	
0.497	65.72	0.05	58.02
.248	74.18	.10	51.89
.199	79.22	.25	47.97
.149	80.57	.50	44.62
		.751	42.63

Johnston, 1906

N	λ	N	λ
99.4°			
0.25	224.0	2	137.8
0.50	194.0	3	108.8
1	174.2	5	85.9

Clausen, 1912

N	λ 6°	λ 18°	λ 30°
0.501	58.57	77.75	98.10
1.000	53.69	70.57	88.37
1.996	46.87	60.57	75.27
3.106	40.05	51.51	63.53

N	$\tau.10^4$ 6°	$\tau.10^4$ 18°	$\tau.10^4$ 30°
0.501	293.4	389.5	491.4
1.000	536.9	705.7	883.7
1.996	935	1209	1502
3.106	1244	1600	1973

Tammann and Rohmann, 1929

P	t	Dλ (%) 1N	Dλ (%) 2N	Dλ (%) 3N
500	20	3.28	2.31	0.92
	40	2.26	1.45	0.50
1000	20	5.37	3.65	1.34
	40	3.57	2.35	0.82
1500	20	6.62	3.84	1.06
	40	4.07	2.64	1.59
2000	20	6.96	3.58	0.39
	40	3.95	2.29	0.03
2500	20	6.74	2.92	-0.97
	40	3.51	1.64	-0.99
3000	20	6.05	1.74	-2.59
	40	2.53	0.28	-2.09

Okazaki, 1935

%	Verdet's constant.10^5 (3441Å)
28°	
6.81	4733
12.22	5000
13.95	5107
20.31	5447
25.14	5695

Mc Clung and Mc Intosh, 1902

d	X-ray absorption (in %)
room temp.	
1.175	96.7
.088	93.0
.044	84.6
.022	74.2
.010	71.3
.000	55.3

Heat constants.

Blümcke, 1884

%	t	U	%	t	U
0	15.0	1	14.85	16.2	0.842
5.12	15.4	0.951	20.23	15.0	.781
9.92	13.0	0.898	23.80	13.7	.754

Faasch, 1911

N	U
18°	
0.422	0.951
0.841	.902
1.692	.812
3.194	.708

Urban, 1932

t	U m = 1.384	U 1.000	U 0.500	U 0.100
10	0.7332	0.7784	0.8758	0.9615
15	.7366	.7875	.8782	.9623
20	.7378	.7892	.8798	.9623
25	.7389	.7900	.8814	.9620
30	.7430	.7906	.8832	.9614
35	.7470	.7914	.8848	.9610
40	.7482	.7922	.8856	.9606

Smith, Stearn and Schneider, 1920

m	mole aq. added	Q dil.(mole aq.)
3.2	35	-2.9
	35	2.7
	25	1.8
2.8	35	0.77
	25	0.88
	15	1.02
1.6	35	0.394
	20	0.41

Lehtonen, 1922

M	Q diss (by 1g chloride)
0°	
0.0625	9.565
.1250	9.341
.2500	8.991
.5000	8.447
1.0000	7.606
1.9998	6.743
2.0000	5.392

Water + Barium bromide ($BaBr_2$)

Heterogeneous equilibria .

Tammann, 1885

%	p	%	p
100°			
10.97	746.3	47.00	610.5
15.84	737.8	51.25	578.6
25.98	713.0	55.81	540.8
34.98	679.8	56.66	534.0
44.78	625.8		

t	p				
	0%	17.25%	28.35%	35.08%	42.34%
40.75	57.6	54.6	51.2	47.3	41.9
45.40	73.4	69.4	65.2	60.8	55.0
48.94	87.8	82.9	78.2	72.8	64.8
51.65	101.4	96.2	89.9	84.4	74.6
55.65	121.3	115.9	107.9	100.6	89.3
58.79	141.3	135.1	125.9	117.3	104.2
61.78	162.2	154.7	144.5	135.2	120.0
65.87	195.2	186.7	174.6	162.5	144.4
69.35	227.3	216.8	202.5	189.3	163.7
72.34	258.5	246.9	231.3	216.5	192.7
75.62	296.9	283.5	264.5	247.8	221.1
78.38	332.8	318.7	298.3	279.7	249.0
80.99	370.0	352.9	329.6	310.1	276.0
81.10	418.8	400.4	375.3	351.3	313.6
87.28	474.0	452.1	424.6	397.8	355.8
89.57	517.2	494.3	464.2	435.2	389.3
91.71	561.1	535.8	503.2	471.3	422.2
93.87	608.0	581.1	544.6	511.5	458.4
94.24	616.6	590.1	552.3	518.8	464.9
96.05	658.9	630.2	590.4	553.0	496.9
97.69	699.4	668.2	626.6	587.9	528.2
100.15	764.0	738.4	685.3	641.3	578.5

Lescoeur, 1894

t	p dissoc. (1+1)	p sat.sol.
10	-	6.2
20	-	10.7
40	23	-
60	58	124
65	68	138
70	91	166
75	115	200
80	-	241

Kremers, 1856 and 1857

Sat. sol. b.t. = 113°

%	f.t.
49.50	0
51.02	20
53.19	40
55.25	60
57.47	80
59.88	100

Johnston, 1905

%	b.t.	%	b.t.
5.71	100.315	39.88	104.71
8.49	100.466	46.81	106.53
13.16	100.735	51.97	108.14
20.94	101.323	56.59	110.02
27.53	102.000	59.03	111.13
32.51	102.619	62.22	111.61
36.50	103.820		

Stokes, 1948

m	osmotic coefficient
25°	
0.1	0.851
.2	.857
.3	.869
.4	.884
.5	.906
.6	.926
.7	.945
.8	.965
.9	.989
1.0	1.013
.2	.063
.4	.112
.6	.162
.8	.212
2.0	.263

Dupuy, 1884

%	f.t.
49.15	15
49.27	30
50.68	60
53.73	100

Etard, 1894

%	f.t.	%	f.t.
45.7	-20	55.5	76
46.5	-9	55.6	77
48.5	+7	56.6	104
48.8	16	60.5	145
49.3	19	59.4	160
50.9	40	60.3	175
55.1	71		

Jones and Chambers, 1900 ,

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	f.t.	M	f.t.
	0°		
0.10	-0.506	0.9032	-7.300
.15	-0.737	1.1290	-7.050
.20	-1.001	.3548	-9.020
.40	-2.039	.5806	-11.260
.4516	-2.310	.8064	-13.860
.50	-2.591	2.2580	-19.030
.6774	-3.890		

Benrath, 1941

%	f.t.
83.0	350
84.0	372
85.0	415

E: 46.6% -22.6 ice + (2+1)

tr.t. 58.5% +113 (2+1)-(1+1)

83% 350 (1+1)- 0 aq.

Properties of phases

Kremers, 1858 and 1859

%	d	%	d
		19.5°	
0	0.998	30	1.327
5	1.043	35	.403
10	.090	40	.483
15	.142	45	.577
20	.199	50	.682
25	.260	55	.797

t		d	
19.5	1.2469	1.4685	1.6611
0	1.2545	1.4809	1.4762
40	.2367	.4553	.6447
60	.2250	.4410	.6283
80	.2116	.4257	.6114
100	.1969	.4097	.5940

Jahn, 1891

c	d
20°	
0	0.9982
34.283	1.2853
64.253	1.5399

Jones and Getman, 1904; Jones, 1905; Jones and Bassett, 1905

M	d	M	d
		0°	
0	0.999868	0.6774	1.174088
0.10	1.025130	0.9032	.231496
.15	.038412	1.1290	.286212
.25	.065340	.3548	.345036
.40	.103536	.5846	.401168
.50	.131416	.8064	.458004
		2.2580	.571168

Heydweiller, 1909

N	d	N	d
		18°	
0.05	1.0051	1.0	1.1248
.10	.0114	2.0	.2486
.20	.0240	4.0	.4913
.50	.0619		

Okazaki, 1935

%	d	%	d
		28°	
7.15	1.0604	29.80	1.3254
12.74	.1162	34.48	.3959
18.71	.1820	39.28	.4752
23.72	.2436		

Thomas and Perman, 1934

%	d	%	d
		30°	
8.74	1.075	31.87	1.341
15.85	.151	42.93	.502
24.19	.243	53.01	.743

%	π	%	π
		0-100 atm.	
8.74	41.2	31.87	32.2
15.85	38.2	42.93	28.3
24.10	35.1	53.01	24.6

Kremers, 1856 and 1857

c	n_D
17°	
0	1.3332
50.6	.3878
102.0	.4322

Jahn, 1891

c	n Hα	n D	n Hβ
	20°		
0	1.3315	1.3332	1.3375
34.283	.3728	.3751	.3805
64.253	.4086	.4113	.4178

Borgesius, 1895

t	$D\ n_D$ (sol.-aq.)
18.5	0.004815
21.2	.004291
21.5	.004300
23.0	.004789
23.0	.001208

Jahn, 1891

c	$(\alpha)_{magn.}$
20°	
0	1
34.283	1.1226
64.253	1.1240

Humburg, 1893

%	$(\alpha)^{mol}_{magn.}$
13.497	18.501

Barbier and Roux, 1890

c	t	dispersive power
13.38	8.7	0.384
26.76	12.4	.423
44.59	12.3	.472
62.43	12.6	.522

Okazaki, 1935

%	Verdet's constant.10^5 (3441 Å)
20°	
7.15	4859
12.74	5233
18.71	5649
23.72	6076
29.80	6645
34.48	7103
39.28	7615

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	λ	M	λ
0°			
0.10	54.25	0.9032	41.80
.15	53.15	1.1290	39.97
.25	50.50	.3548	38.05
.40	48.80	.5806	34.85
.4516	44.70	.8064	28.10
.50	47.11		
.6774	44.38		

Johnston, 1905

N	λ	N	λ
99.4°			
0.001	328.3	1	161.5
.100	225.3	2	149.0
.250	200.3	4	97.6
.500	195.5	8	55.1

Heydweiller, 1909

N	λ	N	λ
18°			
0.05	98.2	1.0	73.9
.10	93.6	2.0	64.7
.20	87.5	4.0	48.3
.50	80.5		

Water + Barium iodide (BaI_2)

Robinson, 1942

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.873	0.9	1.124
0.2	0.897	1.0	1.161
0.3	0.926	1.2	1.238
0.4	0.956	1.4	1.320
0.5	0.988	1.6	1.402
0.6	1.020	1.8	1.483
0.7	1.053	2.0	1.573
0.8	1.089		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.869	0.9	1.122
0.2	0.891	1.0	1.159
0.3	0.918	1.2	1.231
0.4	0.949	1.4	1.308
0.5	0.985	1.6	1.388
0.6	1.017	1.8	1.470
0.7	1.050	2.0	1.599
0.8	1.085		

Lescoeur, 1894

t	p dissoc.
(1+1)	
5	2.75
20	8.35
60	58
80	147
90	200

Etard, 1894

%	f.t.	%	f.t.
58.8	-22	71.9	67
60.0	- 9	72.1	87
60.9	- 5	72.7	96
66.0	+15	73.6	113
66.3	+18	73.7	140
68.3	+32	74.1	165
69.8	+40		

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	f.t.	M	f.t.
0.076	-0.374	1.222	-9.220
.153	-0.764	.528	-13.250
.306	-1.584	.834	-17.500
.612	-3.720	2.139	-24.000
.917	-6.130		

Packer and Rivett, 1926

%	f.t.	%	f.t.	
7.00	-1.0	61.45	-5.8	
12.4	1.7	62.5	0.0	
21.35	3.7	63.6	+5.0	
26.05	5.35	64.8	10.0	
32.05	7.75	65.75	15.0	
35.15	9.35	67.15	19.9	
39.30	12.35	68.8	25.0	
42.95	14.95	68.9	25.7	tr.t.(15+1)
46.1	18.15	68.9	26.0	-(2+1)
50.4	22.9	69.1	30.0	
51.5	24.6	69.6	40.0	
54.6	30.2	70.1	50.0	
56.0 E	33.5	70.7	60.0	
56.8 (15+1)	29.75	71.15	69.5	
57.5	25.4	72.55	87.7	
58.6	20.0	73.35	98.9	tr.t.(2+1)-
59.65	15.4	74.0 (1+1)	110.0	(1+1)
60.65	-9.7	74.3	120.0	

Kremers, 1858 and 1860

%	d	%	d
	19.5°		
0	0.998	35	1.410
5	1.043	40	.493
10	.089	45	.593
15	.141	50	.701
20	.199	55	.822
25	.263	60	.967
30	.331		

t	d			
19.5	1.3673	1.6226	1.8495	2.0453
0	1.3775	1.6380	1.8684	-
40	.3546	.6056	.8294	2.0228
60	.3406	.5884	.8095	2.0011
80	.3254	.5704	.7894	1.9794
100	.3090	.5521	.7695	1.9584

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	d	M	d
	0°		
0.076	1.024336	1.222	1.400456
.153	.050300	.528	.500508
.306	.099552	.834	.600268
.612	.200936	2.139	.699044
.917	.300240		

Heydweiller, 1909

N	d	N	d
	18°		
0.05	1.00696	0.5	1.0808
.10	.01515	1.0	.1624
.20	.0316	2.0	.3237

Packer and Rivett, 1926

%	t	d (sat.sol.)
7.00	- 1.0	1.060
12.40	- 1.7	1:117
21.35	- 3.7	1.221
26.05	- 5.35	1.281
32.05	- 7.75	1.371
35.15	- 9.35	1.423
39.30	-12.35	1.507
42.95	-14.95	1.568
46.1	-18.15	1.634
50.4	-22.9	1.731
51.5	-24.6	1.764
54.6	-30.2	1.842
56.8	-29.75	1.905
57.5	-25.4	1.927
58.6	-20.0	1.952
59.65	-15.4	1.983
60.65	- 9.7	2.012
61.45	- 5.8	2.036
62.5	0.0	2.071
63.6	+ 5.0	2.105
64.8	+10.0	2.144
65.75	+15.0	2.176
67.15	+19.9	2.222
68.8	+25.0	2.277
68.9	+26.0	2.282
69.1	+30.0	2.287
69.6	+40.0	2.304
70.1	+50.0	2.32
70.7	+60.0	2.331
71.15	+69.5	2.35

Kremers, 1857

%	t	n_D
0	17	1.3332
42.96	18	.4219
60.51	17	.4916

Jones and Getman, 1904; Jones , 1904; Jones and Bassett, 1905

M	molecular conductivity
0°	
0.076	112.45
0.153	108.50
0.306	105.72
0.612	101.52
0.917	92.66
1.222	87.36
1.528	79.34
1.834	73.10
2.139	64.40
2.707	50.10
3.125	39.60

Heydweiller, 1909

N	λ
18°	
0.05	102.4
0.10	96.8
0.20	91.3
0.5	82.8
1.0	77.4
2.0	68.4

Water + Barium fluoborate (BaB_2F_8)

Tammann, 1885

%	p
	100°
14.02	738.7
23.70	718.0
36.01	677.9
44.46	641.4
50.38	606.2

Water + Barium thiocyanate ($BaC_2N_2S_2$)

Gibson, 1935

%	π (1-1000 bars)	d
	25°	
0.00	39.35	0.9971
5.62	37.67	1.0358
16.29	34.47	.1231
22.21	32.46	.1779
28.38	30.45	.2427
39.08	26.72	.3648
51.85	22.77	.5427
58.35	20.85	.6497

Wernicke, 1912

t	η(water=1)
	2 M
25	1.106
75.2	1.131

Water + Barium oxide (BaO)

Rosensthiehl and Rühlmann, 1870

%	f.t.	%	f.t.
1.48	0	14.89	59
1.77	6	19.22	64
2.34	12	24.11	69
3.57	22	24.18	70
4.76	30	29.03	73
7.08	41	41.18	77
12.28	54		

Water + Barium hydroxide (BaH_2O_2)

Tammann, 1885

%	p
	100°
7.30	748.6
11.58	741.7
16.01	735.1
24.62	721.9

Lescoeur, 1890

t	p dissoc. (2+1)	p dissoc. (9+1)	p sat.sol.
10	-	2.3	8.3
15	-	3	-
20	-	4.2	-
25	-	6.5	-
30	-	11.5	-
35	-	14	-
40	-	17	-
58	-	84	-
70	-	174	-
74.5	-	218	-
77	14	-	-
100	45	-	-

Haff, 1902

%	f.t.	%	f.t.
1.75	0	20.6	63.4
2.25	5.0	22.4	65.3
3.0	9.5	24.0	66.7
4.8	20.0	25.6	68.1
7.0	30.0	27.1	69.2
8.2	34.5	28.6	70.3
10.4	42.7	31.3	72.2
12.0	47.6	32.7	73.0
14.8	55.0	40.7	76.0
16.9	58.8	46.7	78.0
18.4	60.6	54.23	80.0

%	d	%	d
	80°		
0	0.972	21.36	1.236
3.34	1.009	22.52	1.249
3.84	1.015	24.67	1.278
4.67	1.022	26.13	1.301
5.02	1.031	26.69	1.312
6.26	1.040	27.88	1.338
7.20	1.049	28.59	1.350
8.62	1.062	30.30	1.368
9.83	1.076	30.84	1.375
11.87	1.100	31.81	1.390
13.07	1.114	32.14	1.400
14.04	1.125	32.55	1.413
14.18	1.129	33.72	1.450
15.43	1.152	33.87	1.458
16.89	1.174	36.60	1.479
18.53	1.195	37.54	1.500
19.17	1.200	38.45	1.514
20.12	1.219		

Water + Barium thiohydrate (BaH_2S_2)

Terres and Brückner, 1920

%	f.t.	%	f.t.
0	0	34.5	40
32.0	-15	36.2	60
32.6	0	39.0	80
32.8	+20	43.7	100

Water + Barium chlorite ($BaCl_2O_4$)

Levi, 1923

f.t.	%	f.t.	%
0	30.5	50	33.4
15	31.0	75	38.6
35	31.6	100	44.7

Levi and Bisi, 1956

%	f.t.	%	f.t.
I		II	
12.96	0 (3.5+1)	15.91	6.5 (3.5+1)
14.70	5	16.37	10 anh.
15.12	6.5	17.98	15
16.65	15 anh.	19.58	19.7
16.95	19.7		
16.99	20		
17.40	25		
17.74	30		
18.63	40		
19.40	50		
20.20	60		
20.95	70		
21.80	80		

t	Q diss.
25	-2280 anhydre
8	-3040 anhydre
8	-8660 hydrate

Water + Barium nitrite (BaN_2O_4)

Vogel, 1903

c	f.t.	d
58	0	1.40
63	20	.45
71	25	.50
82	30	.52
97	35	.61

Oswald, 1914

%	f.t.	%	f.t.
9.2	-1.7	58.6	61
19.5	-3.2	61.7	71
33.1	-5.8	67.3	80
34.5	-6.5 E	71.7	92
34.9	-4.3(1+1)	82	110
40.3	+20	92	184
50.3	43	100	217

40% (satd) d^{17}= 1.4895

Bureau, 1937

%	f.t.	E	%	f.t.	E
7.70	-1.4	-5.8	52.0	40.0	-5.85
17.05	-3.05	-5.85	60.8	58.5	"
26.8	-5.25	-5.9	75.1	98.2	"
29.35	-5.95	-5.85	81.25	114.5	-5.9
31.8	+1.0	-5.85	81.75	116	"
34.3	6.65	-5.9	84.2	"	"
39.8	16.0	-5.85	92.8	"	"
49.9	35.2	-5.9			

%	t	d	%	t	d
31.8	1	1.359	51.6	40.4	1.608
41.7	21	.490	58.12	56.0	1.799
47.9	32.1	.579	67.9	79.5	2.097
50.3	37	.605	75.1	98.2	2.265

Water + Barium chlorate ($BaCl_2O_6$)

Tammann, 1885

%	p	%	p
	100°		
7.88	751.3	43.53	668.1
17.52	737.6	45.20	660.9
25.30	722.3	46.32	656.8
30.48	711.0		

Kremers, 1856

%	f.t.	%	f.t.
18.59	0	43.67	60
27.03	20	49.50	80
34.25	40	55.86	100

b.t.= 111°

Tilden and Shenstone, 1885

%	f.t.
66.16	116
74.20	135
78.52	146
83.95	180

Etard, 1894

% (0 -100°) =19.0 +0.382 t

Trautz and Auschütz, 1906

%	f.t.	%	f.t.	
15.28	-2.749	40.05	60	
16.90	0	43.04	70	
21.23	+10	45.90	80	
25.26	20	48.70	90	
27.53	25	51.17	99.1	
29.43	30	52.67	104.6	b.t.(740mm)
33.16	40			
36.69	50			

Carlson, 1910

%	f.t.	%	f.t.
20.32	0	47.95	80
28.15	20	52.89	100
35.84	40	54.55	105.6 b.t.
42.53	60		

Cavallaro, 1941

m	f.t.	m	f.t.
0.01032	-0.05206	0.13514	-0.62800
.02815	.13638	.19152	-0.88846
.05808	.27495	.30040	-1.4059
.07678	.36094	.44337	-2.1286
.10128	-0.47297		

Traube, 1895

%	d
15°	
0	0.99913
1.532	1.01177
3.936	1.03191
6.519	1.05419
10.125	1.08673
14.824	1.13155

Carlson, 1910

%	t	d
20.32	0	1.195
28.15	20	1.274
35.84	40	1.355
42.53	60	1.433
47.95	80	1.508
52.89	100	1.580
54.55	105.6	1.600

Rubien, 1911

N	d
18°	
0	0.99862
0.0975	1.01050
0.1951	1.02232
0.468	1.05511
0.929	1.11010
1.831	1.21566

Clausen, 1912

N	d 6°	d 18°	d 30°
0.2053	1.02543	1.02355	1.02028
.467	.05750	.05499	.05121
.929	.11353	.10999	.10537
1.832	.22120	.21591	.20991

Heydweiller, 1912

N	d 18°
0.1	1.01080
.2	.02293
.5	.05878
1.0	.11865
2.0	.23561

Heydweiller, 1909

N	n_D 18°
0	1.33327
0.5	.34097
1.0	.34831
2.0	.36283

Dn $(H\gamma - H\alpha).10^3 = 23$

Willard and Kassner, 1926

81.472% f.t.= 25° (3+1) d_{25}=2.7753

Rubien, 1911

N	n_D 18°
0	1.33327
0.0975	.33479
.1951	.33626
.468	.34039
.929	.34733
1.931	.36020

Clausen, 1912

N	λ 6°	λ 18°	λ 30°
0.2053	55.19	73.73	93.68
.467	49.38	65.51	83.14
.929	43.46	57.46	72.43
1.832	35.45	46.52	57.90

Heydweiller, 1912

N	κ 18°
0.1	79.6
.2	147.4
.5	323.6
1.0	562.1
2.0	895.6

Water + Barium nitrate (BaN_2O_6)

Tammann, 1885

%	p	%	p
	100°		
5.51	754.0	15.45	741.2
6.64	752.8	19.03	736.0
9.63	749.1	21.43	732.4
10.26	748.1	24.42	727.6

Kremers, 1856

b.t. sat.sol. = 102.5°

Gerlach, 1886

%	b.t.
0	100
11.10	100.5
20.63	101
21.57	101.1

Buchanan, 1899

%	b.t.	%	b.t.
757.56mm		620.6mm	
12.38	100.51	8.21	94.80
14.86	100.61	11.07	94.90
16.68	100.71	13.05	95.00
20.70	100.91	15.32	95.10
22.55	101.01	17.45	95.20
24.68	101.13	19.56	95.30
25.84	101.19	21.53	95.40
		23.59	95.50
		24.23	95.54

Mulder, 1864

%	f.t.	%	f.t.
4.94	0	14.09	49
5.48	4.25	18.17	68
5.84	7	24.30	95
6.54	9	25.82	101.9 b.t.
10.23	31		

Etard, 1894

%	f.t.	%	f.t.
4.3	0.4	16.1	60
4.9	2.1	19.4	73
5.6	6.0	23.4	92
5.6	6.5	27.4	110
6.4	11	31.8	132
7.1	15.3	32.5	134
7.7	18	34.9	150
9.7	28.5	35.4	152
12.8	45.5	38.3	171
14.9	52	45.8	215

Chernai, 1912

%	f.t.	%	f.t.
4.94	0	12.43	40
6.54	10	14.60	50
8.43	20	16.87	60
10.39	30	19.09	70

Sieverts and Petzold, 1933

%		f.t.
2.56	ice	- 0.3
4.06	II	- 4.0
4.28		- 2.4
4.56	E II+ice	- 0.55
4.72	I	0.0
5.13		+ 2.8
5.40		+ 4.6
5.74		+ 6.4
6.16		+ 8.8
6.75		+12.0
7.51		+16.0
8.28		+20.0
12.6		+40.0
16.9		+60.0
21.4		+80.0
25.5		+100.0

Friend and Wheat, 1933

%	f.t.	%	f.t.
4.64	0.2	8.39	20.8
5.69	6.2	11.21	34.2
5.73	7.0	13.56	45.2
6.22	9.1	16.59	58.6
6.25	9.6	21.29	79.4
6.23	9.6	22.12	83.0
6.63	12.0	23.10	86.5
7.02	13.6	23.51	89.6
7.51	15.6	24.80	95.0

Benrath, Gjedebo, Schiffers and Wunderlich, 1937

%	f.t.	%	f.t.	%	f.t.
28.4	112	48.6	226	67.1	334
32.0	135	49.5	234	70.0	348
33.7	144	50.6	239	72.3	362
36.4	158	53.0	256	76.1	380
38.6	171	56.6	276	78.5	393
42.1	191	57.7	281	82.3	417
45.5	209	61.0	297	100	593
45.6	210	64.8	320		

Kasankin, 1891

t	d
sat.sol.	
16	1.0625
32	1.0728

t	height of meniscus in mm
sat.sol.	
17	25.26
34	24.68

Chernai, 1912

%	t	d	%	t	d
4.94	0	1.043	12.43	40	1.104
6.54	10	1.056	14.60	50	1.121
8.43	20	1.073	16.87	60	1.137
10.39	30	1.087	19.09	70	1.146

Borgesius, 1895

%	t	D n_D (sol.-aq.)
7.51	19.3	0.000912
24.17	19.0	0.003589

Water + Barium perchlorate ($BaCl_2O_8$)

Carlson, 1910

% (3+1)	f.t.	% (3+1)	f.t.
67.32	0	83.2	80
74.29	20	84.9	100
78.20	40	86.6	120
81.0	60	88.3	140

Lilich and Dzhurinskii, 1956

m	f.t.
4.11	0
5.29	20
6.13	40

Carlson, 1910

% (3+1)	t	d
67.32	0	1.782
74.29	20	1.912
78.2	40	2.009
81.0	60	2.070
83.2	80	2.114
84.9	100	2.155
86.6	120	2.195
88.3	140	2.230

Kohner, 1928

m	d	n_D
	25°	
0	0.99707	1.33253
0.50015	1.11970	1.34480
0.917445	1.21343	1.35395
1.99864	1.24481	1.37382
2.997725	1.58745	1.38863
3.99986	1.72585	1.40100
4.552605	1.79367	1.40710

Lilich and Mogilev, 1956

m	pH	m	pH
	25°		
0	6.8-6.95	1.255	5.61
0.0558	6.15	1.915	5.17
0.478	5.85	4.581	4.85

Water + Barium sulfamate ($BaH_4N_2O_6S_2$)

King and Hooper, 1941

%	f.t.	%	f.t.
15.31	0.0	33.06	60.0
18.07	10.0	35.60	70.0
20.74	20.0	36.84	75.0
24.43	30.0	39.02	85.0
27.68	40.0	42.31	100.2
30.18	50.0		

Schmelzle and Westfall, 1944

N	d	η(water=1)	N	d	η(water=1)
		25°			
0.0588	1.0052	1.0116	0.7622	1.0931	1.1501
0.1012	.0107	.0186	0.9654	.1179	.2024
0.1980	.0220	.0370	1.2334	.1299	.2334
0.3992	.0480	.0788	1.5850	.1936	.3775
0.5866	.0715	.1166			

Water + Barium acid sulfate ($BaH_2O_8S_2$)

Gillespie and Wasif, 1953

M		M	
0.0688	173.3	0.2599	365.8
0.1371	258.0	0.2988	394.1
0.3865	436.0	0.8448	524.7
0.4760	469.3	0.9513°	526.0
0.6096	501.6	1.042	523.4
0.6640	515.2	1.122	519.3
0.7147	516.0	1.200	513.5

Water + Barium dithionate (BaS_2O_6)

Tammann, 1885

%	p	%	p
		100°	
21.82	746.2	33.66	730.0
22.10	746.0	34.89	729.1
30.57	734.8		

De Baat, 1926

%	f.t.	%	f.t.
7.56	0	15.75	20
12.45	12	19.86	30

Ishikawa and Kimura, 1926

%	f.t.	d
7.75	0	1.0702
9.68	5	1.0877
11.56	10	1.1049
13.61	15	1.1248
15.63	20	1.1430
17.67	25	1.1631
19.70	30	1.1839
21.61	35	1.2010
23.55	40	1.2207
25.18	45	1.2390
26.91	50	1.2560
28.54	55	1.274
30.13	60	1.291
31.57	65	1.305
33.08	70	1.324
34.25	75	1.337
35.69	80	1.355
37.20	85	1.374
38.42	90	1.387

Water + Barium formate ($BaH_2C_2O_4$)

Krasnicki, 1887

%	f.t.	%	f.t.
21.73	0	27.08	50
22.16	10	28.04	60
23.14	20	28.54	70
24.44	30	28.42	80
25.82	40	(2+1)	

Stanley, 1904

%	f.t.	%	f.t.
23.24	0	28.22	73.6
23.32	10	31.63	93.2
24.42	35	32.84	100
26.67	54.5		

Ashton, Houston and Saylor, 1933

%	f.t.	%	f.t.
20.68	0	27.72	60
21.88	10	29.27	70
23.05	20	30.61	80
24.20	30	32.19	90
25.37	40	33.86	100
26.51	50		

de Garcia, 1920

N	d	n_D	N	d	n_D
		23°			
1	1.092120	1.3466	0.125	1.010788	1.3345
0.5	.045998	.3396	0.062	.004857	.3336
0.25	.022676	.3362			

Heydweiller, 1921

N	d	λ	N	d	λ
0.2	1.01856	63.6	1	1.0907	43.44
0.5	1.04588	53.3	2	1.1787	31.19

Gladstone and Hibbert, 1897

%	molecular refraction (formate) H_α	D	H_β
		18°	
26.0	41.10	41.53	42.22

Water + Barium acetate ($BaC_4H_6O_4$)

Tammann, 1885

%	p
100°	
17.81	735.7
29.41	713.9
36.16	700.7
41.35	688.9

Stokes, 1953

Isopiestic solutions

m_1	m_2	m_1	m_2
0.1073	0.1384	1.283	1.755
.1510	.1964	.464	1.973
.1902	.2500	.667	2.220
.2593	.3434	.870	2.454
.5045	.6880	2.062	2.649
.6496	.8926	2.822	3.332
1.087	1.493	3.474	3.812

m_1 = molality of barium acetate

m_2 = molality of sodium chloride

Tilden and Shenstone, 1884

%	f.t.	%	f.t.
32.66	22	46.12	130
43.34	40	47.89	136
44.13	60	58.60	180
44.22	110		

Krasnicki, 1887

%	f.t.	%	f.t.
36.69	0	43.61	50
39.02	10	43.81	60
40.81	20	43.65	70
42.15	30	43.17	80
43.06	40		

(1+1)

Walker and Fyffe, 1903

%	f.t.	%	f.t.
37.03(3+1)	0.3	44.13 0aq.	40.5
38.12	7.9	44.04	41.5
40.90	17.5	43.79	44.5
42.13	21.6	43.34	51.8
43.85	24.1	42.73	63.0
43.31(2+1)	26.2	42.37	73.0
42.89	30.6	42.54	84.0
43.12	35.0	42.79	99.2
43.79	39.6		

Franz, 1872

%	d	%	d
17.5°			
0	0.9987	21	1.1593
1	1.0074	22	.1679
2	.0161	23	.1764
3	.0248	24	.1850
4	.0335	25	.1936
5	.0423	26	.2026
6	.0487	27	.2110
7	.0550	28	.2206
8	.0614	29	.2296
9	.0678	30	.2386
10	.0744	31	.2495
11	.0816	32	.2605
12	.0888	33	.2715
13	.0960	34	.2825
14	.1032	35	.2937
15	.1106	36	.3058
16	.1187	37	.3179
17	.1268	38	.3300
18	.1348	39	.3421
19	.1425	40	.3540
20	.1507		

Heydweiller, 1912

N	d	N	d
18°			
0.1009	1.00808	0.995	1.08823
.2011	.01714	2.165	.18975
.503	.04456	4.35	.36755

Guillaume, 1946

%	t	d
14.42	20	1.1136
31.5	21	1.2722

Heydweiller, 1913

N	n_D
18°	
0	1.33327
0.2	.33658
0.5	.34136
1.0	.34909
2.0	.36397

Guillaume, 1946

%	t	n 5780Å	$(\alpha)^*_{magn.}10^6$
14.42	20	1.3540	3.708
31.5	21	1.3797	3.370

* in radians, gauss, centim.

Mc Gregory, 1894

N	$\varkappa$	N	$\varkappa$
18°			
0.0001	0.0819	0.01	7.237
.0002	.1637	.05	31.24
.0006	.4840	.10	57.31
.0010	.7974	.50	210.2
.0020	1.566	1.00	331.3
.0060	4.487	2.00	448.6

Heydweiller, 1912

N	$\varkappa$
18°	
0.1009	61.1
0.2011	107.9
0.503	219.1
0.995	342.2
2.165	452.2
4.350	328.3

Water + Barium propionate ($BaC_6H_{10}O_4$)

Tammann, 1885

%	p
100°	
24.06	729.6
39.90	700.0

Krasnicki, 1887

%	f.t.	%	f.t.
32.42	0	38.55	50
34.01	10	39.30	60
35.40	20	39.93	70
36.62	30	40.43	80
37.66	40		

Wing and Thompson, 1926

%	f.t.	%	f.t.
36.5	0.3	38.9	65.3
36.2	5.1	39.8	75.6
36.0	15.0	40.2	80.4
36.2	24.8	41.2	85.6
36.6	34.8	42.2	90.5
37.1	44.8	43.2	95.4
38.0	55.0	44.7	100.7

Heydweiller, 1921

N	d	λ
18°		
0.2	1.01881	51.7
0.5	.04657	41.21
1	.0922	31.41
1.5	.1367	25.79
2	.1810	18.79

Water + Barium butyrate ($BaC_8H_{14}O_4$)

Deszathy, 1893

%	f.t.	%	f.t.
27.23	0.0	26.71	50.0
26.82	10.0	27.38	60.0
26.54	20.0	28.35	70.0
26.39	30.0	29.64	80.0
26.36	40.0		

(2+1)

Water + Barium valerate ($BaC_{10}H_{18}O_4$)

Fürth, 1888

%	f.t.	%	f.t.
17.82	0	16.60	50
17.10	10	17.07	60
16.61	20	17.79	70
16.07	30	18.75	80
16.06	40		

Water + Barium trimethylacetate ($BaC_{10}H_{18}O_4$)

Landau, 1893

%	f.t.	%	f.t.
25.54	0	24.24	50
25.27	10	24.81	60
24.86	20	26.02	70
24.52	30	27.92	80
24.15	40		

(5+1)

Water + Barium caproate ($BaC_{12}H_{22}O_4$)

Keppich, 1888

%	f.t.	%	f.t.
8.68	0	8.01	50
8.04	10	8.62	60
7.66	20	9.48	70
7.52	30	10.55	80
7.63	40		

(2+1)

Water + Barium methyl-3-valerate ($BaC_{12}H_{22}O_4$)

Kulisch, 1893

%	f.t.	%	f.t.
10.48	0	6.21	50
7.73	10	7.74	60
6.44	20	9.98	70
5.55	30	12.82	80
5.48	40	16.17	90

(7+2)

Water + Barium methyl-2-valerate ($BaC_{12}H_{22}O_4$)

König, 1894

%	f.t.	%	f.t.
12.54	0	11.37	50
11.76	10	11.99	60
11.24	20	12.80	70
11.00	30	13.97	80
11.04	40	15.21	90

(4+1)

Water + Barium benzenesulfonate ($BaC_{12}H_{10}O_6S_2$)

Tammann, 1885

%	p
100°	
19.51	744.5
27.92	734.5
31.43	728.3

Water + Barium phenolsulfonate ($BaC_{12}H_{10}O_8S_2$)

Tammann, 1885

%	p
100°	
16.38	748.3
28.30	736.7
46.37	704.4
52.81	687.7

LII. WATER + SALTS OF Zn, Cd and Hg .

Water + Zinc chloride ($ZnCl_2$)

Tammann, 1885

%	p
100°	
14.58	734.5
27.12	693.4
40.78	604.1
48.59	515.9
54.71	421.3

Menzies and Bovina, 1912

%	p 14.64°	p 24.64°	p 29.60°
60	4.4	8.1	10.8
65	2.5	5.0	6.9
70	1.4	2.9	4.1

Lescoeur, 1894

t	p dissoc.	t	p dissoc.
hydrate			
20	2	130	118
90	22	140	145
100	35	150	193
110	57	160	258
120	86		

Robinson and Stokes, 1940

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1085	0.1502	0.6323	0.8724
0.1086	0.1494	0.6681	0.9179
0.1161	0.1608	0.8884	1.206
0.1431	0.1982	0.9078	1.223
0.1571	0.2181	1.036	1.380
0.1935	0.2703	1.251	1.634
0.1986	0.2762	1.440	1.856
0.2480	0.3466	1.632	2.096
0.3383	0.4728	2.045	2.622
0.3855	0.5407	2.149	2.752
0.4098	0.5743	2.182	2.807
0.4881	0.6815	2.539	3.306
0.5116	0.7182	3.035	4.028
0.5828	0.8141	3.187	4.291

m_1 = m of $ZnCl_2$ m_2 = m of KCl

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
25°			
0.1	0.847	3.0	.858
.2	.845	3.5	.903
.3	.842	4.0	.955
.4	.838	4.5	1.022
.5	.833	5.0	.091
.6	.829	6.0	.229
.7	.824	7.0	.361
.8	.817	8.0	.500
.9	.810	9.0	.628
1.0	.805	10.0	.753
.2	.792	12.0	.961
.4	.782	14.0	2.09
.6	.781	16.0	.18
.8	.785	18.0	.22
2.0	.792	20.0	.24
2.5	.820	22.0	.25

Chufarov, Zhuravleva and al., 1953

%	b.t.	%	b.t.
61	120	85.5	180
66.6	130	87.4	190
72.9	140	92.9	210
77.1	150	95.3	220
80.2	160	97.2	230
83.0	170		

Etard, 1894

%	f.t.	%	f.t.
54.8	-20	59.1	+5
55.4	-14	60.2	+9
56.5	-10	62.0	+15
57.4	-4	66.8	+33
57.9	-1	68.3	+42

Dietz, 1899

%	f.t.	%	f.t.	%	f.t.
(3+1)		(5+2)		(2+1)	
64.5	-5	67.42	0	67.56	0
67.58	0	71.96	8	73.70	10
71.57	+7	75.14	13	79.07	19
(3+2)		(1+1)		anh.	
67.45	0	74.33	0	79.12	15
73.65	10	78.25	11	81.19	20
80.08	20	84.61	27	82.21	41
83.43	26			83.51	60
				86.01	100

Jones and Getman, 1904 and Jones, 1904

M	f.t.	M	f.t.
0.0493	-0.263	0.394	-2.098
.0986	0.509	0.592	3.221
.197	1.020	1.787	10.850
.296	1.543	3.574	25.500

Mylius and Dietz, 1905

%	f.t.
12.3	-5
20.0	-10
45.3	-40
51.0	-62 E
53.0	-50 (4+1)
55.9	-40
61.5	-30 tr.t.
65.4	-10 (3+1)
67.5	0
69.7	+5
71.6	6.5
73.8	5
70.1	0 (5+2)
71.6	6.5 (3+1)-(5+2)
73.1	10.0 (5+2)
75.2	12.5 "
77.0	11.5 (5+2)-(3+2)
78.3	9.0 (5+2)-(1+1)
79.4	6.0 (5+2)
74.9	-6.0 (3+2)
75.5	0.0 (3+1)-(3+2)
76.8	+10.0 (3+2)
78.6	20.0 "
80.9	26.0 "
81.2	26.3 "
77.4	0 (1+1)
78.4	10 "
79.8	20 "
81.3	28 (1+1) - anh.
82.7	31 (1+1)
81.2	25 anh.
81.9	40 "
83.0	60 "
84.4	80 "
86.0	100 "
100.0	262 "

Frankenheim, 1847

t	d	t	d
0	1.3631	55	1.3189
5	.3594	60	.3139
10	.3555	65	.3098
15	.3517	70	.3050
20	.3477	75	.3004
25	.3437	80	.2959
30	.3396	85	.2912
35	.3358	90	.2864
40	.3313	95	.2817
45	.3270	100	.2768
50	.3227	105	.2721

Kremers, 1858

%	d	%	d
	19.5°		
1	1.008	31	1.302
2	.018	32	.314
3	.027	33	.327
4	.035	34	.338
5	.043	35	.350
6	.051	36	.364
7	.061	37	.378
8	.070	38	.390
9	.080	39	.404
10	.089	40	.418
11	.098	41	.430
12	.108	42	.444
13	.118	43	.458
14	.126	44	.471
15	.135	45	.486
16	.144	46	.497
17	.163	47	.515
18	.163	48	.530
19	.173	49	.547
20	.184	50	.563
21	.194	51	.578
22	.205	52	.597
23	.216	53	.612
24	.226	54	.628
25	.236	55	.647
26	.247	56	.666
27	.258	57	.683
28	.268	58	.701
29	.279	59	.721
30	.289	60	.737

t	d			
	1*	2	3	4
0	1.1335	1.2558	1.3991	1.5697
19.5	.1256	.2440	.3846	.5525
40	.1143	.2312	.3676	.5332
60	.1008	.2135	.3498	.5137
80	.0970	.2090	.3453	.5091
100	.0694	.1786	.3122	.4735

concentration from table of density

Fouqué, 1867

%	d	
	0°	11.4°
3.58	1.0836	1.0820
17.10	.4694	.4629
24.44	.8753	.8645

van der Willigen, 1869

%	d
25°	
23.00	1.21011
31.05	.30021
31.50	.30553
35.98	.36070

Long, 1880

%	d	%	d
15°			
2.5	1.024	30	1.299
5	.048	40	.423
10	.094	50	.570
20	.190	60	.746

Blümcke, 1884

%	t	d
4.7	15.3	1.041
24.1	15	.229
38.0	17	.393
58.5	17	.709
68.0	16.5	.899

Lussana, 1897

%	t	d
0	23	0.9976 (?)
4.53	21.53	1.0016
7.47	22.70	.0045
18.93	23.30	.0184
42.72	23.12	.0590

Chéneveau, 1907

%	d
16°	
0	0.9990
9.39	1.0882
17.53	.1657
24.66	.2390
30.95	.3072

Heydweiller, 1912

N	d	N	d
18°			
0.0976	1.00478	1.024	1.05905
.1978	.01103	1.927	.10821
.4981	.02920	3.836	.20509

Rabinowitch, 1921

%	d	%	d
25°			
		36.75	1.3722
3.08	1.0250	44.20	.4752
6.14	.0534	49.55	.5590
12.22	.1103	56.22	.5826
18.36	.1686	61.72	.7802
22.17	.2067	68.43	.9230
28.71	.2782	74.87	2.0741

Okazaki, 1935

%	d	%	d
28°			
7.97	1.0679	37.15	1.3739
8.21	.0701	38.22	.3877
17.29	.1550	46.19	.5008
23.33	.2157	55.16	.6510
29.42	.2808	55.20	.6520
34.55	.3408		

Wagner, 1883

%	η(water°=1) 15°	25°	35°	45°
33.752	1.5170	1.1790	0.8997	0.7264
23.487	1.1150	0.8662	.6977	.5749
15.334	0.9361	0.7274	.5782	.4821

Rabinowitsch, 1921

%	η(water=1)	%	η(water=1)
25°			
0	1.000	36.75	2.50
3.08	.108	44.20	3.41
6.14	.202	49.55	4.51
12.22	.372	56.22	6.91
18.36	.546	61.72	12.7
22.17	.670	68.43	33.7
28.71	2.000	74.87	153

Lussana, 1897

P	t	R	P	t	R
	42.72%			23.362%	
150	9.14	784.1	150	8.30	1961.1
300	9.27	741.8	300	8.20	1898.0
450	9.57	706.1	450	8.27	1856.3
600	10.41	663.8	600	8.78	1804.5
750	11.16	628.9	750	9.80	1741.1
880	12.00	598.6	880	9.92	1720.1
1000	12.76	573.9	1000	10.52	1678.8
150	20.59	600.3	150	21.87	1377.5
300	20.74	581.8	300	21.85	1358.2
450	20.58	566.8	450	22.16	1329.3
600	20.48	552.1	600	22.40	1303.3
750	21.08	533.6	750	22.11	1294.7
880	20.57	525.3	880	22.13	1281.7
1000	20.67	513.7	1000	22.08	1273.9
150	38.32	492.1	150	37.54	1075.9
300	38.74	479.3	300	37.78	1060.8
450	38.93	467.9	450	38.05	1045.0
600	39.46	455.1	600	38.09	1033.0
750	40.02	434.3	750	38.50	1016.6
880	40.02	434.3	880	38.57	1006.1
1000	40.80	422.8	1000	38.78	996.6
150	54.37	450.7	150	56.75	851.6
300	55.40	440.1	300	57.58	835.1
450	56.02	430.1	450	57.98	821.5
600	56.84	420.1	600	58.11	813.5
750	57.78	409.9	750	59.13	796.1
880	58.91	400.9	880	58.79	791.7
1000	59.79	392.7	1000	58.94	783.9
	7.47%			4.51%	
150	7.20	5568.5	150	4.53	13128
300	6.91	5462.8	300	4.52	12923
450	6.94	5342.0	450	4.76	12669
600	7.04	5237.5	600	4.59	12596
750	7.25	5143.7	750	5.04	12263
880	7.29	5189.8	880	5.13	12120
1000	7.67	5009.6	1000	5.27	12981
150	22.97	3657.3	150	22.39	8159
300	23.02	3602.0	300	22.49	8050
450	22.54	3589.4	450	22.39	7980
600	22.91	3521.6	600	22.48	7875
750	22.72	3494.5	750	22.19	7848
880	22.62	3470.9	880	22.01	7811
1000	22.59	3451.3	1000	21.91	7774
150	41.27	2705.6	150	39.45	5474
300	40.99	2688.6	300	39.45	5420
450	41.63	2635.3	450	39.25	5376
600	40.92	2641.7	600	39.39	5306
750	40.12	2656.6	750	38.93	5287
880	40.49	2619.1	880	38.81	5244
1000	39.47	2646.5	1000	38.51	5247
150	59.71	2094.5	150	58.84	3874
300	61.48	2022.0	300	59.01	3824
450	63.91	1934.5	450	59.00	3776
600	64.75	1892.6	600	59.20	3728
750	63.47	1909.3	750	58.63	3622
880	63.12	1903.5	880	57.94	3723
1000	63.25	1889.1	1000	57.89	3698

*R = Specific resistance in ohms .

Lilich and Varshavski, 1956

m	pH	m	pH
	25°		
0	6.8-6.95	4.88	4.02
0.035	6.45	5.98	3.61
.125	5.95	9.13	2.92
.397	5.40	12.07	2.50
1.350	4.75	15.64	1.93
1.646	4.70	18.62	1.50
2.780	4.35	23.30	0.50
4.10	4.05		

Godlewski, 1902

%	e	%	e
	25°		
0.0068	0	31.03	0.2928
.034	0.0602	35.48	.3024
.068	.0857	39.32	.3126
.11	.1017	42.86	.3228
.20	.1207	45.49	.3297
.34	.1366	50.00	.3411
.67	.1584	52.51	.3522
.99	.1704	54.55	.3606
1.48	.1829	56.52	.3642
2.44	.1986	59.06	.3806
3.29	.2081	60.78	.3902
4.76	.2200	62.50	.4050
6.39	.2303	63.64	.4174
12.12	.2528	64.61	.4301
27.10	.2839		

Mc Clung and Mc Intosh, 1902

d	% X-ray absorption
room temp.	
1.198	94.1
.104	90.5
.053	83.1
.026	73.2
.012	65.2
.000	58.9

Sresnievski, 1881

%	t	a^2
0	23.02	15.39
16.56	25.56	14.90
30.96	23.48	11.88
45.30	-	7.17

Kremers, 1859

$d^{19.5}$	n_D
15°	
0.9983	1.3334
1.1986	.3730
.3667	.4037
.5495	.4359

van der Willigen, 1869

spectrum lines	n_D				
	25°				
	0%	23%	31.05%	31.50%	35.98%
A	1.36252	1.37071	1.38654	1.38755	1.39732
a	.36344	.37161	.38758	.38855	.39843
B	.36422	.37243	.38844	.38940	.39926
C	.36502	.37325	.38931	.39030	.40019
D	.36719	.37548	.39169	.39273	.40264
F	.36985	.37822	.39464	.39575	.40574
b	.37034	.37875	.39523	.39630	.40632
F	.37213	.38059	.39721	.39830	.40839
G	.37477	.38302	.39984	.40091	.41117
G	.37628	.38498	.40195	.40297	.41339
H	.37705	.38580	.40284	.40387	.41438
H	.37834	.38716	.40428	.40534	.41590
H	.37981	.38878	.40601	.40716	.41780

Chêneveau, 1907

%	n_D
16°	
0	1.3334
9.39	.3525
17.53	.3674
24.66	.3810
30.95	.3939

Long, 1880

%	κ	τ.10⁴	%	κ	τ.10⁴
		15°			
2.5	262	213	30	880	172
5	459	192	40	802	198
10	691	165	50	598	232
20	867	156	60	350	307

Heydweiller, 1912

N	κ	N	κ
	18°		
0.0976	83.0	1.024	557.6
.1978	154.3	1.927	768.6
.4981	335.6	3.836	907.8

Rabinowitsch, 1921

%	κ	%	κ
	25°		
3.08	369	36.75	985
6.14	617	44.20	854
12.22	884	49.55	732
18.36	980	56.22	535
22.17	1019	61.72	365
28.71	1022	68.43	182
		74.87	88

Lussana, 1897

t	R*	t	R
42.72%		23.36%	
11.21	778.8	10.77	1910.0
13.65	723.3	11.45	1878.4
20.29	623.3	21.62	1408.2
20.48	621.2	21.72	1404.4
37.54	506.4	36.44	1106.8
41.36	485.9	38.78	1071.6
53.19	427.0	55.07	877.6
60.19	395.1	58.29	847.8
8.074%		4.53%	
7.60	5662.5	5.29	13016
8.20	5581.5	5.30	13004
21.96	3798.5	21.42	8437
22.71	3738.5	22.30	8260
40.50	2774.3	38.10	5683
40.52	2774.0	38.70	5615
57.58	2186.0	58.19	3960
64.59	1991.0	58.50	3939

Da Silveria, 1939

Raman effect (see author)

Okazaki, 1935

%	Verdet's constant.10^5 (3441 Å)	%	Verdet's constant.10^5 (3441 Å)
7.97	4893	37.15	7149
8.21	4943	38.22	7176
17.29	5550	46.19	8004
23.33	5951	46.62	8073
29.42	6443	55.16	9023
34.55	6924	55.20	9065

Marignac, 1876

%	U	%	U
	19.51°		
3.64	0.9590	23.23	0.7960
7.03	.9330	33.52	.7042
13.14	.8842	43.06	.6212

Blümcke, 1884

%	U	%	U
	room temp.		
4.7	0.972	58.5	0.514
24.1	.805	68.0	.437
38.0	.685		

Beetz, 1879

d	c.10^5	t
1.132	436	8-14
	660	28-36
1.310	410	8-14
	656	28-36
1.870	370	8-14
	445	28-36

$$c = \text{heat conductivity} = \frac{(\lg t_0 - \lg t)}{t}$$

Jäger, 1891

%	relative heat conductivity coefficient
0	100
17.5	91.5
35	83.7

Water + Zinc bromide ($ZnBr_2$)

Lescoeur, 1894

t	p	t	p
	sat.sol.		
20	2	100	74
40	3	110	98
90	52	120	120

Ishikawa and Yoshida, 1932

t	p sat.sol.	p diss. (2+1)-anh.
20	1.73	1.30
25	2.15	1.75
30	2.59	2.33
35	3.15	3.15
40	4.22 (anh.)	
45	5.57	
50	7.35	

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.869	3.0	1.066
0.2	0.886	3.5	1.100
0.3	0.911	4.0	1.143
0.4	0.937	4.5	1.195
0.5	0.962	5.0	1.253
0.6	0.984	6.0	1.379
0.7	1.002	7.0	1.521
0.8	1.018	8.0	1.667
0.9	1.032	9.0	1.809
1.0	1.039	10.0	1.941
1.2	1.047	12.0	2.15
1.4	1.049	14.0	2.28
1.6	1.047	16.0	2.34
1.8	1.043	18.0	2.35
2.0	1.042	20.0	2.33
2.5	1.048		

Dupuy, 1884

%	f.t.	%	f.t.
22.10	15	66.10	60
24.60	30	68.54	100

Etard, 1894

%	f.t.	%	f.t.
66.3	-20	83.8	107
68.8	+4	85.0	170
77.5	22	89.3	210
83.6	97	100.0	375

Dietz, 1899

%	f.t.	%	f.t.
	(3+1)		
77.13	-15	80.64	- 5
78.45	-10		
	(2+1)		
79.06	- 8	82.46	+25
77.55	0	84.08	+30
80.76	+13	86.20	+37
	anhydre		
85.45	+35	86.57	+ 80
85.53	+40	87.05	+100
86.08	+60		

Kremers, 1858 - 1859

%	d	%	d
	19.5°		
0	0.998	35	1.398
5	1.043	40	1.473
10	1.091	45	1.557
15	1.194	50	1.647
20	1.202	55	1.752
25	1.263	60	1.872
30	1.328		

t	d					
0	1.18982	1.36158	1.53566	1.72403	1.87035	2.12379
19.5	.18290	.34960	.52000	.70530	.84940	.09920
40	.17228	.33449	.50156	.68394	.82605	.07234
60	.15911	.31769	.48204	.66195	.80218	.04556
80	.14377	.29906	.46095	.63867	.77731	.01815
100	.12645	.27893	.43885	.61451	.75177	1.99059

Kremers, 1859

%	d
19.5°	
0	0.9983
30.50	1.3349
52.97	1.7161
67.85	2.1060

Rabinowitsch, 1921

%	d	%	d
	25°		
1.06	1.0057	41.73	1.4847
5.66	.0482	51.82	.6634
13.20	.1233	61.65	.8844
19.14	.1852	71.37	2.1670
26.04	.2670	77.50	.3865
33.49	.3622	83.67	.6492

%	η(water=1)	%	η(water=1)
	25°		
1.06	1.029	41.73	2.08
5.66	.098	51.82	2.78
13.20	.234	61.65	4.26
19.14	.364	71.37	8.83
26.04	.525	77.50	19.5
33.49	.730	83.67	78.8

Kremers, 1859

%	t	n_D
0	16	1.3333
30.50	16	.3349
52.97	13	.7161
67.85	18	2.1060

Rabinowitsch, 1921

%	κ	%	κ
	25°		
1.06	94	41.73	1100
5.66	438	51.82	921
13.20	889	61.65	692
19.14	1096	71.37	405
26.04	1201	77.50	223
33.49	1192	83.67	96

Lilich and Varshavskii, 1956

m	pH	m	pH
	25°		
0.024	6.72	7.07	2.92
1.78	4.48	8.23	2.64
2.53	4.37	12.00	1.60
3.09	4.20	13.59	0.94
5.64	3.45	16.70	0.05

Water + Zinc iodide (ZnI_2)

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.893	2.0	1.282
0.2	0.924	2.5	1.262
0.3	0.957	3.0	1.262
0.4	0.994	3.5	1.278
0.5	1.038	4.0	1.297
0.6	1.083	4.5	1.335
0.7	1.124	5.0	1.381
0.8	1.163	6.0	1.487
0.9	1.194	7.0	1.617
1.0	1.220	8.0	1.766
1.2	1.260	9.0	1.911
1.4	1.283	10.0	2.034
1.6	1.291	12.0	2.205
1.8	1.292		

Etard, 1894

%	f.t.	%	f.t.
70.9	-18	82.1	97
74.0	-5	83.0	100
80.4	+17	83.5	107
80.3	47	83.6	138
81.3	62	83.8	140
81.2	73		

Dietz, 1899

%	f.t.	%	f.t.
(2+1)		anhydre	
80.50	-10	81.11	0
80.77	- 5	81.20	+18
81.16	0	81.66	+40
82.06	+10	82.37	+60
83.12	+22	83.05	+80
89.52	+27	83.62	100

Kremers, 1858 and 1860

%	d	%	d	%	d
		19.5°			
0	0.998	30	1.366	55	1.751
5	1.043	35	1.388	60	1.872
10	1.089	40	1.418	65	2.017
15	1.138	45	1.557	70	2.176
20	1.194	50	1.647	75	2.356
25	1.253				

t	d				
0	1.23939	1.52403	1.80371	1.99442	2.42424
19.5	.23200	.50960	.78410	.97130	.3936
40	1.22129	1.49233	1.76159	1.94538	2.36074
60	.20832	.47334	.73791	.91865	.32806
80	.19323	.45274	.71283	.89099	.29560
100	.17612	.43060	.68684	.86283	.26332

Rabinowitsch, 1936

%	d	%	d
	25°		
0.978	1.0050	38.94	1.4620
5.33	.0442	49.20	.6984
11.85	.1074	58.70	.8676
17.73	.1707	68.37	2.1587
24.47	.2531	73.57	.3550
31.13	.3420	81.49	.7301

%	η(water=1)	%	η(water=1)
	25°		
0.978	1.011	38.94	1.69
5.33	.048	49.20	2.22
11.85	.112	58.70	3.08
17.73	.188	68.37	5.17
24.47	.298	73.57	7.82
31.13	.453	81.49	23

%	κ	%	κ
	25°		
0.978	65	38.94	1284
5.33	318	49.20	1032
11.85	670	58.70	736
17.73	962	68.37	453
24.47	1225	73.57	315
31.13	1341	81.49	128

Kremers, 1859

$d^{19.5}$	n_D
17°	
0.9983	1.3332
1.4010	.4049
1.5547	.4308
2.0252	.5090

Lilich and Varshavski, 1956

m	pH	m	pH
	25°		
0.156	5.88	5.14	3.25
1.29	4.80	6.81	2.27
2.34	4.40	7.79	1.78
3.39	4.28	7.90	1.75
4.91	3.48	8.85	1.03

Water + Zinc chlorate ($ZnCl_2O_6$)

Meusser, 1902

%	f.t.	%	f.t.	%	f.t.
(6+1)		(4+1)		anh.	
55.62	-18	66.52	18	30.27	-13
59.19	0	67.66	30	26.54	-9
60.20	8	69.06	40		
67.32	15	75.44	55		

sat.sol. d^{18}= 1.916

Traube, 1895

%	d
15°	
0	0.99862
2.692	1.02083
5.462	1.04391
8.181	1.06729
12.341	1.10510
16.927	1.14845

Rubien, 1911

N	d	N	d
18°			
0	0.99862	1.0570	1.09312
0.1051	1.00826	2.0977	.18443
.21085	.01782	4.203	.36377
.5266	.04621		

Heydweiller, 1912

N	d
18°	
0.0985	1.00763
0.1970	1.01658
0.4925	1.04320
0.985	1.0870
1.970	1.1731
3.940	1.3416

Heydweiller, 1909

N	n_D	N	n_D
18°			
0	1.33327	1.0	1.34730
0.1	.33476	2.0	.36083
.2	.33618	4.0	.38616
.5	.34043		

Rubien, 1911

N	n_D	N	n_D
18°			
0	1.33327	1.0570	1.34807
0.1051	.33484	2.0977	.36211
.21085	.33632	4.203	.38861
.5266	.34079		

Heydweiller, 1912

N	$\varkappa$	N	$\varkappa$
18°			
0.0985	73.0	0.985	537.9
.1970	136.3	1.970	873.4
.4925	305.4	3.940	1139

Water + Zinc bromate ($ZnBr_2O_6$)

Heydweiller, 1921

N	d	λ
18°		
0.5	1.0686	51.4
1	.1360	43.30
2	.2691	32.34
3	.4004	24.22

Water + Zinc nitrate (ZnN_2O_6)

Tammann, 1885

%	p	%	p
	100°		
8.52	743.7	29.70	615.5
17.59	715.0	36.99	595.5
24.54	682.0	43.84	528.5

Dieterici, 1923

m	p	m	p
	0°		
0	4.579	1.875	4.077
0.498	.482	2.997	3.590
0.994	.363	3.718	3.236
1.409	.217		

Ewing and Fisher, 1937

%	p				
	20°	30°	40°	50°	60°
10.258	16.78	30.41	53.04	88.97	-
18.466	15.99	29.24	51.00	85.62	-
25.577	15.60	28.31	49.70	82.55	-
26.434	15.12	27.31	48.34	80.24	-
32.354	14.25	25.31	45.21	75.01	-
34.618	13.62	25.17	44.07	72.60	-
35.079	13.37	23.33	42.42	71.63	-
44.045	11.05	13.98	35.19	59.49	-
53.236	8.04	14.85	26.25	43.16	-
57.362	6.30	11.45	20.32	35.55	-
63.150	-	7.76	13.91	24.65	40.96
67.514	-	4.75	8.74	15.17	26.21
72.713	-	2.46	8.42	8.42	14.74

%	p					
	35°	40°	45°	50°	55°	60°
80.177	1.77	2.38	3.20	4.21	5.65	7.11
82.375	1.02	1.45	1.87	2.48	3.28	4.33

Ewing and Fischer, 1937

t	p dissoc.	t	p dissoc.
(6+1)		(2+1) I	
10	4.25	37	1.80
20	7.27	38	2.00
25	9.01	39	2.15
30	10.76	40	2.24
33	11.45	42	2.36
34	11.55	45	2.48
35	11.42	50	2.44
36	10.55	55	2.06
36	9.66		
35	8.15	(2+1) II	
34	7.54		
(4+1)		35	1.57
		30	1.15
34	7.54	25	0.55
35	7.95	20	0.42
36	8.36		
38	8.74	(4+1)-(2+1)	
40	9.18		
42	9.14	39	2.42
43	9.02	38	2.25
44	8.56	37	1.80
44	8.73	35	1.25
43	4.21	30	0.65
41	3.43	25	0.27
39	2.42	20	0.08
38	2.25		
(6+1)-(4+1)			
34	7.54		
30	5.95		
25	4.15		
20	2.92		
15	2.00		
10	1.36		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.862	1.6	1.238
0.2	0.873	1.8	1.296
0.3	0.890	2.0	1.355
0.4	0.909	2.5	1.510
0.5	0.934	3.0	1.664
0.6	0.958	3.5	1.814
0.7	0.982	4.0	1.957
0.8	1.009	4.5	2.098
0.9	1.037	5.0	2.235
1.0	1.064	6.0	2.489
1.2	1.120	6.5	2.600
1.4	1.180		

Rüdorff, 1872

%	f.t.
6.36	-1.80
7.84	-2.35
11.18	-3.55
15.02	-5.35
18.41	-7.35

Dupuy, 1884

%	f.t.
74.84	15
75.10	30
87.66	60

Funk, 1899

%	f.t.	%	f.t.
anh.		(6+1)	
30.00	-16	44.63	-18
35.00	-23.5	45.26	-15
39.65	-29.5	45.51	-13
42.85	-34	45.75	-12
		48.66	0
(9+1)		52.00	+12.5
		53.50	18
40.12	-25	55.90	25
40.75	-22.5	63.63	36.4
42.03	-20	64.73	36
43.59	-18	65.83	33.5
		66.38 (3+1)	37
		67.42	40
		68.21	41
		69.26	43
		77.77	45.5

Jones, 1904 and Jones and Getman, 1904

M	f.t.	M	f.t.
0.065	-0.322	1.290	-8.930
.129	-0.633	.548	-11.800
.258	-1.281	.806	-14.720
.516	-2.812	2.064	-18.240
		2.580	-27.000

Vasilyev, 1909

mol%	f.t.
15.4	35.4 (6+1) - (4+1)
20.0	45.5 (4+1)
24.2	38.4 (4+1) - (2+1)
33.3	55.0 (2+1)

Ewing, Mc Govern and Mathews, 1933

%	f.t.
56.1	25.1(6+1)
56.3	25.9
56.9	27.3
57.6	28.8
57.7	29.1
58.2	29.7
58.3	30.4
58.9	31.4
59.1	31.4
59.2	31.6
60.3	32.4
60.3	33.3
61.0	34.1
63.4	36.1
63.7	35.9
65.2	35.4
67.4	35.1
66.2	34.6(6+1)-(4+1)
67.9	40.0(4+1)
70.0	43.2
72.5	44.7
75.8	42.4
76.3	41.3
77.2	39.7
77.3	39.1
77.8	38.2
77.8	37.5
79.0	32.2
78.0	37.2(4+1)-(2+1)
79.7	43.6(2+1)
79.8	44.9
80.1	46.6
81.6	50.6
81.9	51.9
84.0	55.4
85.2	53.8
86.0	53.2
86.3	52.1(2+1)-(1+1)
86.6	52.8(1+1)
86.6	54.1
87.6	59.2
88.6	65.2
89.0	66.2
89.4	68.6
89.8	70.3
90.0	70.7

Sieverts and Petzold, 1933

%	f.t.	
16.4	-5.8	
25.4	-12.0	
31.7	-19.6	
37.1	-28.0	
38.9	-32.0	E
39.7	-29.0	(3+1)
40.8	-23.0	
42.9	-19.5	
43.5	-23.0	(6+1)
44.5	-19.0	
48.3	+0.4	
58.1	30.0	
61.2	35.0	
65.0	35.6	(6+1)-(4+1)
69.7	43.5	(4+1)
70.2	45.0	
73.8	45.0	
75.8	43.5	
77.9	37.0	(4+1)-(2+1)
80.7	51.0	(2+1)
83.2	54.0	
86.2	51.8	(2+1)-(1+1)
87.2	59.0	(1+1)
89.9	73.1	
91.2	73.9	
92.6	73.0	

Oudemans, Jr., 1868

%	d	%	d
	14°		
0	0.9993	15.91	1.1450
0.64	1.0045	16.55	.1515
1.27	.0097	17.18	.1581
1.91	.0150	17.82	.1647
2.55	.0204	18.45	.1714
3.18	.0258	19.09	.1782
3.82	.0312	19.73	.1850
4.45	.0367	20.36	.1919
5.09	.0423	21.00	.1989
5.73	.0479	21.63	.2060
6.36	.0536	22.27	.2131
7.00	.0593	22.91	.2203
7.63	.0650	23.54	.2276
8.27	.0708	24.18	.2349
8.91	.0767	24.82	.2422
9.54	.0826	25.46	.2496
10.17	.0886	26.09	–
10.81	.0947	26.73	.2647
11.45	.1007	27.37	.2724
12.09	.1069	28.00	.2802
12.73	.1131	28.64	.2880
13.37	.1194	29.27	.2960
14.00	.1257	29.91	.3040
14.64	.1321	30.55	.3120
15.27	.1385	31.18	.3201
		31.82	.3292

Franz, 1872

%	d	%	d	%	d
		17.5°			
1	1.0086	18	1.1790	35	1.3888
2	.0185	19	.1900	36	.4021
3	.0284	20	.2008	37	.4154
4	.0383	21	.2131	38	.4287
5	.0483	22	.2254	39	.4420
6	.0576	23	.2379	40	.4553
7	.0670	24	.2499	41	.4688
8	.0764	25	.2623	42	.4825
9	.0858	26	.2741	43	.4962
10	.0954	27	.2875	44	.5099
11	.1056	28	.3001	45	.5239
12	.1158	29	.3127	46	.5384
13	.1260	30	.3251	47	.5528
14	.1362	31	.3379	48	.5673
15	.1472	32	.3506	49	.5818
16	.1570	33	.3634	50	.5964
17	.1680	34	.3762		

Barnes and Scott, 1898

%	d	%	d
	17.3°		
0	0.9987	14.39	1.1284
1.210	1.0087	19.65	1.1830
1.574	1.0118	29.21	1.2933
5.923	1.0491	30.86	1.3136
7.091	1.0597	41.32	1.4579
11.36	1.0988	47.28	1.5504

Jones, 1904 and Jones and Getman, 1904

M	d	M	d
	0°		
0.065	1.008904	1.548	1.233788
0.129	1.019796	1.806	1.270492
0.258	1.039380	2.064	1.310168
0.516	1.081908	2.580	1.380700
1.290	1.196804		

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	1.0111	1.07602
0.1026	1.00666	2.0147	1.15095
0.2048	1.01460	4.0042	1.29527
0.5076	1.03773		

Heydweiller, 1912

N	d	N	d
	18°		
0.1040	1.0067	1.040	1.0781
.2080	.0148	2.080	.1554
.5200	.0387	4.160	.3057

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
15.700	1.1210
15.909	1.1223
31.405	1.2406
31.818	1.2438

Okazaki, 1935

%	d
28°	
9.00	1.0738
11.88	.1007
18.69	.1698
26.22	.2541
32.62	.3346
39.61	.4319
45.72	.5251

Guillaume, 1946

%	d
20°	
10.95	1.0968
19.61	.1843
26.33	.2610
33.77	.3538

Wagner, 1883

%	η (water°=1)			
	15°	25°	35°	45°
44.500	1.6790	1.3060	1.0540	0.8791
30.626	1.0472	0.8572	0.6952	0.5769
15.955	0.8073	0.6428	0.5260	0.4376

Bak, 1946

%	mol%	η			
		80°	60°	40°	20°
4.03	0.4	330	540	730	1100
10.6	1.11	430	590	800	1290
29.3	3.79	540	780	1090	3150
41.0	6.18	880	1130	3060	7150
49.9	8.68	1180	2120	5110	10450
54.4	10.20	1740	2990	10960	-
61.1	13.10	3010	5490	11100	-
63.7	14.30	3670	6680	11600	-
66.1	15.02	4340	7130	-	-
71.8	19.55	6350	11220	-	-
76.0	23.20	8740	15310	-	-

Morgan and Schrainen, 1913

%	σ	%	σ
45°			
0	68.46	31.57	74.34
10.09	69.64	34.69	75.18
14.15	70.29	38.47	76.62
17.40	71.01	41.29	77.66
19.42	71.27	44.19	78.59
21.21	71.80	46.18	79.36
23.68	72.16	49.44	80.61
26.70	72.88	54.41	82.17
29.38	73.45		

Heydweiller, 1909

N	n_D	N	n_D
18°			
0	1.33327	1.0	1.34675
0.1	.33466	2.0	.35971
.2	.33603	4.0	.38415
.5	.34008		

Rubien, 1911

N	n_D	N	n_D
18°			
0	1.33327	1.0111	1.34690
0.1026	.33470	2.0147	.35991
.2048	.33609	4.0042	.38420
.5076	.34018		

de Mallemann and Guillaume, 1945

(d^{20}= 1.2588) $(\alpha)_{magn.10^5}$= 18.92

Guillaume, 1946

%	n 5780 Å	(α)*magn.10^6
20°		
10.95	1.3518	3.652
19.61	.3670	3.394
26.33	.3798	3.190
33.77	.3955	2.972

*in radians, gauss, centim.

Jones, 1904 and Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.065	93.40	1.548	52.43
0.129	90.56	1.806	46.53
0.258	82.50	2.064	42.24
0.516	73.00	2.580	36.03
1.290	53.89	2.632	27.09

Heydweiller, 1912

N	κ	N	κ
	18°		
0.1040	83.6	1.040	610.5
.2080	155.9	2.080	980.0
.520	347.4	4.160	1245.0

Bak, 1946

%	κ 20°	κ 40°	κ 60°	κ 80°
5	40	55	75	95
10	85	120	150	200
30	140	190	235	280
40	115	160	220	270
50	90	135	180	241
55	75	120	160	225
60	50	75	125	180
63.7	40	60	110	175
65	50	65	90	125
70	35	60	60	90
77	30	45	53	70
77.8	-	40	50	65
80	-	50	60	75

Okazaki, 1935

%	Verdet's constant.10^5 (3514 Å)
28°	
9.00	4183
11.88	4191
18.69	4198
26.22	4198
32.62	4237
39.61	4214
45.72	4248

Lilich and Varshavskii, 1956

m	pH	m	pH
	25°		
0.022	5.85	1.88	3.55
.155	4.95	2.38	3.24
.408	4.54	3.93	2.63
.613	4.42	5.22	2.12
.920	4.15		

Da Silveira, 1939

Raman effect(see author)

Marignac, 1876

%	U	%	U
	20-52°		
5.00	0.9461	29.60	0.7176
9.52	.8990	41.20	.6410
17.38	.8234	51.24	.5906

Water + Zinc perchlorate ($ZnCl_2O_8$)

Lilich and Dzhurinskii, 1956

m	f.t.	m	f.t.
3.97	0	4.3	25
4.01	5	4.37	30
4.07	10	4.46	35
4.11	15	4.52	40
4.17	20	4.62	45
		4.74	50

Stokes, 1948

m	Osmotic coefficient	m	Osmotic coefficient
	25°		
0.1	0.893	1.2	1.450
0.2	0.928	1.4	1.578
0.3	0.966	1.6	1.708
0.4	1.010	1.8	1.843
0.5	1.056	2.0	1.986
0.6	1.105	2.5	2.358
0.7	1.157	3.0	2.739
0.8	1.212	3.5	3.117
0.9	1.269	4.0	3.494
1.0	1.328	4.5	3.790

Lilich and Mogilev, 1956

m	pH	m	pH
	25°		
0	6.8-6.95	1.086	1.65
0.101	3.03	.76	.20
.137	2.85	.85	.26
.263	2.56	2.79	0.64
.520	2.15		

Lilich and Varshavskii, 1956

m	pH	m	pH
	25°		
0.052	4.07	1.11	2.10
.198	3.22	2.32	1.12
.556	2.50	3.50	0.17
1.15	2.13	4.91	-0.35

Da Silveira, 1939

Raman effect (see author)

Water + Zinc sulfate ($ZnSO_4$)

Heterogeneous equilibria.

Tammann, 1885

t	p 0%	p 23.75%	p 35.05%
41.17	58.9	56.5	53.6
43.32	65.9	63.4	59.3
49.19	88.9	84.9	79.8
52.29	103.6	99.2	93.5
57.73	140.9	135.3	127.5
60.83	155.3	148.9	140.2
66.43	200.0	192.4	182.2
70.70	241.0	232.4	220.2
71.00	244.1	235.6	223.3
74.38	281.9	271.5	258.2
76.16	303.7	293.2	278.9
79.10	342.7	330.9	315.3
82.97	400.5	386.5	368.6
84.78	430.1	416.0	396.3
87.22	473.0	457.8	436.8
90.05	526.9	510.3	488.1
92.39	575.5	556.5	531.7
94.39	620.0	599.7	574.7
97.10	684.6	663.8	636.2
100.22	766.4	743.5	714.4

%	p
100°	
13.48	750.0
21.03	744.2
26.34	734.6
30.38	725.0
43.73	673.0

Emden, 1887

t	p	t	p	t	p
12.93%		21.01%		21.04%	
20.27	17.3	19.79	16.5	20.24	16.95
25.44	23.7	25.58	23.5	25.88	23.8
30.55	32.0	30.62	31.5	31.84	33.95
35.00	40.95	35.44	41.1	35.04	40.65
39.67	52.7	39.98	52.6	39.45	51.5
45.31	70.9	45.46	70.15	43.97	65.35
51.66	97.55	50.13	89.0	50.22	89.7
55.46	117.3	50.71	91.5	54.58	111.1
61.05	152.3	54.70	110.95	59.41	139.3
65.06	182.6	59.98	142.9	63.81	170.2
70.91	235.55	65.83	185.3	70.33	226.4
76.12	294.5	69.24	215.3	75.94	288.4
80.45	353.0	76.08	289.65	80.48	348.45
85.45	430.4	80.66	350.5	87.73	430.6
90.59	524.8			90.16	512.3
95.22	625.3			95.83	634.25

Schüller, 1889-90

t	p 0%	p 14.3%	p 25.0%
51.35	100.765	99.53	97.95
55.15	118.343	116.76	114.88
64.70	184.476	182.11	179.48
64.95	186.533	184.16	182.18
69.80	231.670	228.81	225.89
74.00	276.624	273.07	270.16
75.00	288.517	284.92	281.66
75.50	294.678	291.03	287.68
77.40	318.808	315.06	311.46
77.60	321.527	317.79	313.79
86.70	462.813	458.63	458.63
91.50	556.111	550.72	550.72
94.70	627.517	620.61	620.61
	33.3%	40.0%	45.4%
51.35	96.16	94.54	91.96
55.15	113.04	110.77	107.94
64.70	176.53	173.36	169.75
64.95	178.51	175.44	171.44
69.80	222.34	218.26	213.55
74.00	266.56	261.48	255.75
75.00	277.71	272.73	266.76
75.50	282.82	278.28	272.63
77.40	306.03	303.30	295.28
77.60	308.76	303.89	298.82
86.70	447.80	439.18	431.06
91.50	537.39	529.26	519.73
94.70	-	596.81	587.16

t	p 0%	p 21.7%	p 35.7%	p 45.4%
13.20	11.311	11.031	10.688	10.202
13.50	11.503	11.210	10.847	10.363
13.60	11.610	11.320	10.941	10.442
15.50	13.117	12.767	12.342	11.770
15.75	13.327	12.987	12.485	11.995
16.00	13.536	13.169	12.755	12.188
16.30	13.802	13.452	12.877	12.445
18.00	15.537	15.152	14.623	13.997
21.00	18.495	17.995	17.383	16.628
25.50	24.269	23.612	22.801	21.847
29.65	30.829	29.987	28.004	27.829
34.70	41.148	41.104	39.857	37.254
35.05	41.945	41.870	39.543	37.950
40.00	54.906	53.542	51.797	49.838
44.85	70.851	69.096	67.009	64.485
49.90	91.533	89.423	86.812	83.545

Dieterici, 1923

m	p	m	p
	0°		
0	4.579	1.489	4.472
0.497	4.552	1.997	4.411
0.992	4.518	2.557	4.304

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.590	1.0	0.478
0.2	0.533	1.2	0.489
0.3	0.506	1.4	0.508
0.4	0.492	1.6	0.533
0.5	0.483	1.8	0.566
0.6	0.476	2.0	0.602
0.7	0.473	2.5	0.717
0.8	0.473	3.0	0.861
0.9	0.474	3.5	1:033

m	activity coefficient	m	activity coefficient
	25°		
0.1	(0.150)	1.0	0.043
0.2	0.104	1.2	0.040
0.3	0.083	1.4	0.038
0.4	0.071	1.6	0.036
0.5	0.063	1.8	0.035
0.6	0.057	2.0	0.035
0.7	0.052	2.5	0.037
0.8	0.048	3.0	0.041
0.9	0.046	3.5	0.048

Wiedemann, 1874

t	p dissoc.	t	p dissoc.
25	14.5	55	80.1
30	19.1	60	100.9
35	28.2	65	121.7
40	40.3	70	145.4
45	52.4	75	186.6
50	56.8	80	230.3

Frowein, 1887

t	p dissoc.	t	p dissoc.
(7+1)-(6+1)		(6+1)-(5+1)	
18.00	8.406	17.85	7.633
20.45	10.075	20.45	9.477
25.15	14.697	25.15	13.286
28.35	19.135	28.90	17.448
29.95	21.389	29.95	18.826
		31.70	21.075

Lescoeur, 1890

t	p dissoc. (7+1)	p dissoc. X aq.
10	4.9	5.2
20	9.1	10.0
30	18.9	20.2
40	-	44.2
50	-	73.1
60	-	113.9
70	-	170.8

Ishikawa and Murooka, 1933

t	p dissoc.
20	15.92 (7+1)
25	21.17
30	27.73
35	35.90
40	45.68

Gerlach, 1886

%	b.t.	%	b.t.
0	100	37.89	103
11.58	100.5	40.48	103.5
20.00	101	42.83	104
27.38	101.5	44.66	104.5
31.22	102	46.15	105
35.02	102.5		

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
2.80	100.08	23.35	100.92
6.23	100.17	24.75	101.02
9.20	100.27	26.03	101.15
11.80	100.38	27.20	101.27
15.04	100.47	28.49	101.42
18.16	100.60	29.23	101.49
20.12	100.71	30.77	101.71
22.03	100.83		

Johnston, 1907

%	b.t.	%	b.t.
1.18	100.074	10.30	100.192
4.17	.091	10.94	.242
6.63	.135	14.06	.297
8.22	.167		

Poggiale, 1844

%	f.t.	%	f.t.
30.08	0	42.60	60
32.60	10	44.21	70
34.69	20	45.83	80
36.87	30	47.31	90
38.84	40	48.73	100
40.75	50		

Tobler, 1855

%	f.t.
29.2	0
34.6	20
35.3	25
40.1	50
44.6	70

Mulder, 1864

%	f.t.	%	f.t.
30.55	0	37.81	25
32.52	8.25	39.43	31
33.38	10	41.59	39
34.89	15.4		

de Coppet, 1872

%	f.t.	%	f.t.
12.95	-1.67	24.93	-5.1
16.03	-2.2	26.53	-6.1
18.71	-2.75	28.00	-7.15
21.04	-3.3	30.54	-10.1

Rüdorff, 1872

%	f.t.	%	f.t.
5.10	-0.60	18.71	-2.80
6.88	-0.80	21.04	-3.65
9.35	-1.10	23.11	-4.35
12.95	-1.70	24.93	-5.11
16.03	-2.15		

Etard, 1894

%	f.t.	%	f.t.
29.1	+1	46.5	77
32.6	13	44.7	100
34.8	20	43.0	111
40.2	41	40.7	125
40.9	49	38.0	137
43.4	55	37.4	144
45.0	62	30.0	169
47.0	70	29.0	171

Barnes, 1900

%	f.t.	%	f.t.
29.43	0	41.70	41.49
33.66	15.00	42.68	46.40
33.85	15.88	43.51	49.97
38.46	30.70	43.41	49.99
41.36	39.92	43.50	50.00
41.37	39.95	43.51	50.02
41.43	40.73		

Kahlenberg, 1901

%	f.t.	%	f.t.
1.59	-0.258	9.85	-1.246
4.78	-0.625	12.03	-1.493
8.23	-1.030	14.48	-1.922

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.051	-0.094	1.015	-1.753
.102	.194	.421	-2.715
.203	.372	.624	-3.327
.406	.697	.827	-3.976
.609	-1.027	2.032	-4.990

Koppel, 1905

E (7+1) -6.5°

Cohen, Chattaway and Tombrock, 1907

$\% = 29.48 + 0.522\ t + 0.00496\ t^2$

Cohen and Sinnige, 1909

P	% electric method	% direct method
	25°	
250	36.75	36.79
500	36.75	36.78
750	36.68	36.71
1000	36.71	36.63

Cohen, 1924 and Cohen and Moesveld, 1925

%	f.t.	%	f.t.
29.39	0	39.82	35
30.75	5	40.52	37
32.09	10	41.03	39
33.48	15	42.21	45
34.44	18	43.25	50
36.49	25	44.42	55
38.00	30		

tr.t.= 37.9° (7+1) - (6+1)

Bury, 1924

%	f.t.	%	f.t.
(6+1)		(7+1)	
46.3	61.1	38.3	22.94
44.9	56.1	38.2	22.88
42.2	44.48	37.9	21.56
40.5	35.03	37.7	20.90
39.2	27.63	37.1	18.23
39.0	25.18	35.5	11.62
38.9	24.90	34.8	7.63
38.3	21.02		
38.0	18.21		
37.4	13.83		
37.1	11.35		

Caven and Johnston, 1926

36.61% f.t.= 25°

Cohen, Inonye and Euwen, 1910

P	tr.t.
(7+1)	
1	38.12
500	39.96
1000	41.19
1500	42.63

Rohmer, 1940

%	f.t.	%	f.t.	%	f.t.
(7+1)		(6+1)		(4+1)	
38.1	30	41.2	40	46.6	65
38.7	32	42.2	45	47.3	70
40.2	36	44.6	56	47.9	75
(2+1)		(1+1)			
44.3	56	42.2	56		
44.2	60-75	41.9	60		
		41.4	65		
		40.8	70		
		40.2	75		
		39.2	85		
		37.7	100		

tr.t.	(7+1) - (6+1)	40.8%	37.9°
	(6+1) - (1+1)	43.0%	48.8°
	(6+1) - (2+1)	44.2%	54.6°
	(6+1) - (4+1)	46.5%	63.4°

Benrath, 1941

%	f.t.	%	f.t.
50	90	20	230
45	115	15	255
40	145	10	250
35	175	5	255
25	205		

Karnaukhov and Shevchuk, 1953 (fig.)

%	f.t. I (7+1)	f.t. II	(6+1)
10	-2	-	-
27	-5.8	-	-
30	+5	-	-
34	20	7.6	-
37	28	21	11.4
38	32	24.9	24.9
39	38.8	-	38.8
42	-	-	48

Hoeltzenbein, 1924

t	c	diffusion ratio
13.1	72.793	0.27982
15.6	31.2	.30058
18.5	10.0	.36905

Properties of phases (density)

Schiff, 1858 and 1859

%	d	%	d
20.5°			
0	0.9981	15.492	1.1719
3.873	1.0378	23.235	.2767
7.746	.0797	34.856	.4622
11.615	.1250		

%	d	%	d	%	d
20.5°					
0.56	1.0038	11.79	1.1267	23.01	1.2731
1.12	.0096	12.35	.1334	23.57	.2810
1.68	.0154	12.91	.1402	24.13	.2893
2.24	.0212	13.47	.1470	24.69	.2976
2.81	.0270	14.03	.1539	25.25	.3059
3.37	.0329	14.59	.1608	25.81	.3143
3.93	.0388	15.15	.1678	26.37	.3228
4.49	.0448	15.71	.1749	26.93	.3314
5.05	.0508	16.27	.1820	27.49	.3410
5.61	.0569	16.83	.1892	28.05	.3486
6.17	.0630	17.40	.1965	28.62	.3574
6.73	.0690	17.96	.2038	29.18	.3663
7.29	.0752	18.52	.2112	29.74	.3754
7.85	.0815	19.08	.2187	30.30	.3846
8.42	.0879	19.64	.2263	30.86	.3938
8.98	.0942	20.20	.2339	31.42	.4031
9.54	.1006	20.76	.2416	31.98	.4125
10.10	.1071	21.32	.2494	32.54	.4225
10.66	.1136	21.88	.2572	33.10	.4316
11.22	.1201	22.44	.2651	33.66	.4413

Gerlach, 1867

%	d	%	d
15°			
0.56	1.005	17.40	1.199
1.12	.012	17.96	.208
1.68	.018	18.52	.215
2.24	.023	19.08	.223
2.81	.027	19.64	.230
3.37	.034	20.20	.239
3.93	.040	20.76	.245
4.48	.046	21.32	.254
5.04	.052	21.88	.262
5.60	.058	22.44	.2698
6.16	.065	23.00	.279
6.72	.072	23.56	.287
7.28	.077	24.12	.294
7.85	.084	24.68	.303
8.42	.089	25.25	.309
8.98	.096	25.82	.319
9.54	.102	26.38	.329
10.10	.109	26.94	.336
10.66	.115	27.48	.345
11.22	.122	28.05	.3520
11.78	.129	28.61	.361
12.34	.136	29.17	.369
12.90	.142	29.73	.379
13.46	.149	30.30	.389
14.03	.150	30.86	.3974
14.59	.163	31.43	.405
15.15	.170	31.99	.415
15.71	.178	32.55	.424
16.27	.184	33.11	.434
16.83	.1923	33.66	.4438

Fouqué, 1867

%	d 4°	d 12°
1.00	1.0052	1.0051
4.52	.0289	.0283
13.96	.0994	.0978

Lundqvist, 1869

%	t	d
0	40.8	0.9919
20.82	40.3	1.227
21.38	48.0	.238
30.07	45.2	.368

Ewing and Mc Gregor, 1872-73

%	d	%	d
	10°		
0	0.9997		
1.37	1.0140	22.46	1.2709
1.81	.0187	23.70	.2891
2.67	.0278	23.78	.2895
5.10	.0540	24.41	.2987
7.02	.0760	26.43	.3288
9.36	.1019	28.07	.3530
14.04	.1582	31.19	.4053
16.04	.1845	31.84	.4174
18.71	.2186	32.04	.4220
21.38	.2562		

Favre and Valson, 1874

m	d	m	d
	23.5°		
0	0.997	4	1.307
1	1.084	5	.376
2	.162	6	.443
3	.236		

Grotrian, 1877

%	t	d	t	d	t	d
7.37	8.13	1.0804	18.51	1.0780	30.81	1.0744
11.06	9.06	.1228	18.80	.1201	28.85	.1176
14.82	7.78	.1692	18.60	.1658	32.62	.1608
					38.03	.1587
19.63	8.15	.2322	18.11	.2285	29.92	.2240
					41.60	.2192
22.56	7.76	.2734	16.98	.2697	38.70	.2606
29.74	7.75	.3823	18.48	.3779	39.00	.3686

Kohlrausch, 1879

%	d	%	d
	18°		
5	1.0509	20	1.2323
10	.1069	25	.3045
15	.1675	30	.3788

Rönnberg, 1880

t	d	t	d
	10 c		20 c
3.7	1.09993	4.0	1.19326
9.7	.09883	10.0	.19168
16.2	.09722	16.2	.18981
21.2	.09582	21.2	.18826
25.6	.09457	25.4	.18676
	30 c		40 c
3.7	1.28140	4.0	1.36489
10.7	.27912	10.2	.36247
15.7	.27695	15.6	.36009
21.2	.27461	21.2	.35747
26.0	.27238	25.5	.35539

The author writes % instead of c .

Fink, 1885

%	d
	18°
0.96	1.0090
9.67	.1028
23.50	.2822
29.64	.3730

Schüller, 1889-90

t	d 21.7%	d 35.7%	d 45.4%
14.4	1.22017	-	-
14.5	-	1.36123	-
14.7	-	-	1.45455
20.6	1.21914	1.36002	.45333
30.4	.21839	.36001	.45241
39.8	.21741	.35783	.45156
49.8	.21730	.35780	.45197 (?)

Charpy, 1893

%	d	%	d
	0°		
0	0.9999	17.7573	1.2143
5.1110	1.0564	21.4444	.2663
9.7426	.1105	24.7170	.3150
14.0307	.1644		

de Lannoy, 1895

t	relative volume			
	4%	10%	20%	30%
0	1.00000	1.00000	1.00000	1.00000
10	.00053	.00103	.00178	.00258
20	.00215	.00326	.00438	.00550
30	.00492	.00610	.00763	.00890
40	.00860	.00976	.01148	.01291
50	.01281	.01408	.01590	.01750
60	.01800	.01902	.02100	.02251
70	.02360	.02461	.02662	.02785
80	.03020	.03103	.03290	.03390

Sentis, 1897

%	t	d
1	16.45	1.0844
2	17.2	.1679
3	17.6	.2482
4	16.55	.3253

Barnes and Scott, 1898

%	d	%	d
	18.5°		
0	0.9985	8.44	1.0894
0.577	1.0045	11.20	.1220
1.46	.0138	17.08	.1957
3.18	.0318	21.28	.2528
3.82	.0387	25.14	.3091
6.65	.0696	29.22	.3718

Pann, 1901

%	d		
	10°	18°	30°
0	0.999727	0.998622	0.995673
10.00	1.10910	1.10695	1.17472
16.17	.18234	.18200	.33914
27.63	.34774	.34360	.50863

Brümmer, 1902

%	d	%	d
	15°		
0	0.9993	8.487	1.0913
2.937	1.0302	11.221	.1236
5.752	1.0609	13.858	.1554

Jones, 1904. Jones and Getman, 1904

M	d	M	d
	0°		
0.051	1.008260	1.015	1.155816
.102	.016388	.421	.215116
.203	.031760	.624	.243636
.406	.063716	.827	.271400
.609	.094456	2.032	.300008

Chéneveau, 1907

%	t	d
0	14.9	0.9992
22.46	14.6	1.2720

Keitanpää and Rentanch, 1911-12

N	d			
	20°	30°	40°	50°
0.1	1.0085	1.0081	1.0078	1.0075
.2	.0162	.0155	.0152	.0150
.5	.0381	.0376	.0372	.0370
.8	.0610	.0603	.0598	.0597
1.0	.0744	.0737	.0731	.0728
2.0	.1455	.1443	.1438	.1433

Herz, 1914

N	d	N	d
	25°		
0	0.9971	4.02	1.3100
1.39	1.1057	5.56	1.4071
2.78	1.2099		

Herz, 1917

N	d	N	d
	25°		
0	0.9971	4.020	1.3101
1.390	1.1060	5.560	1.4070
2.780	1.2100		

Stocker, 1920

%	t	d
0	22.7	0.9977
7.49	21.1	1.0786
11.11	21.1	.1204
15.41	20.9	.1743
22.36	21.9	.2678

Suominen, 1922

t	d	t	d	t	d
2N		1.5N		1N	
0.0	1.1518	0.0	1.1173	0.0	1.0815
10.3	.1500	9.8	.1156	10.4	.0799
21.0	.1464	23.0	.1117	20.5	.0775
32.1	.1422	31.9	.1085	33.3	.0729
41.0	.1372	42.6	.1034	40.1	.0698
51.1	.1330	50.0	.0999	50.5	.0657
59.5	.1271	60.0	.0942	60.2	.0599
73.5	.1180	76.0	.0842	72.4	.0523
81.4	.1129	87.0	.0769	82.2	.0452
0.5N		0.25N		0.1N	
0.0	1.0412	0.0	1.0242	0.0	1.0098
9.9	.0404	10.5	.0235	9.8	.0094
23.0	.0377	21.7	.0212	21.8	.0073
31.9	.0349	30.6	.0189	32.0	.0045
41.7	.0308	39.5	.0159	40.6	.0008
50.0	.0270	51.3	.0103	51.8	0.9969
61.5	.0205	61.8	.0043	61.2	.9910
73.5	.0132	80.3	0.9952	74.0	.9836
89.0	.0007	89.0	0.9856	86.2	.9723

Cohen, Helderman and Moesveld, 1924

%	d	%	d
	30°		
0	0.99556	17.14	1.08113
2.15	1.01669	25.66	1.07354
4.30	1.04031	34.21	1.06688
8.60	1.08744		

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
14.771	1.1359
15.337	1.1403
29.542	1.2666
29.705	1.2699

Geffcken, 1929

m	d	m	d
	25°		
0	0.99707	1.486	1.22391
0.4989	1.07657	1.5046	.22660
0.9901	.15153	2.0200	.29827
1.0020	.15335	2.0214	.29844

Gibson, 1934

mol%	d	mol%	d
	25°		
0	0.9930	17.82	1.2031
4.97	1.0493	25.04	.3046
11.48	1.1234	27.08	.3356

Guillaume, 1946

%	d	%	d
	20°		
9.81	1.1070	19.90	.2350
12.08	.1332	29.95	.3832

Schmidt, 1859

%	t	π
7.5	16.3	45.1
13.5	16.7	41.0
17.5	16.0	41.6
30	13.2	34.3
32.5	16.6	32.3
41	17.5	26.4

Gibson, 1934

mol %	π (1-1000 bars)	mol %	π (1-1000 bars)
	25°		
0	39.46	17.82	23.89
4.97	34.67	25.04	18.79
11.48	29.07	27.08	17.42

Viscosity and surface tension

Grotrian, 1876

%	t	η	t	η
1.44	9.46	1932	13.77	1663
3.67	9.50	2135	14.31	1840
7.41	9.55	2591	14.42	2216
11.08	9.12	3266	13.56	2763
14.85	9.22	3201	13.90	3548
19.61	15.23	4201	21.77	3935
22.61	9.31	8115	14.29	6662
29.75	15.34	14160	21.43	10810

Grotrian, 1877

%	t	η	t	η	t	η
7.37	10.84	1742	19.88	1368	30.45	1075
11.06	10.58	2086	20.22	1616	30.21	1277
14.82	9.57	2656	25.40	1768	39.80	1279
19.63	9.77	3685	25.03	2423	39.78	1700
22.56	8.78	4718	19.87	3397	30.61	2610
29.74	8.30	9117	20.31	5985	31.08	4402

t	η	t	η
22.56%		29.74%	
40.73	2006	39.93	3476
48.61	1749	48.70	2862

Wagner, 1883

%	η (water°=1) 15°	25°	35°	45°
7.12	0.9706	0.7926	0.6266	0.5153
16.64	1.5603	1.1861	0.9420	0.7346
23.09	2.3282	1.7741	1.3532	1.0811

Keitanpää and Rentanch, 1911-12

N	η(water^t=1) 20°	30°	40°	50°
0.1	1.0349	1.0334	1.0299	1.0267
.2	.0637	.0599	.0565	.0547
.5	.1704	.1629	.1566	.1488
.8	.2580	.2510	.2435	.2318
1.0	.3259	.3191	.3067	.2985
2.0	.7745	.7507	.7332	.6966

Herz, 1917

N	η	N	η
25°			
0	895	4.020	3753
1.390	1406	5.560	6522
2.780	2252		

Guillaume, 1946

%	$(\alpha)^*_{magn.}10^6$ 5780 Å	n
20°		
0	3.974	-
9.81	3.719	1.3529
12.08	.656	.3562
19.9	.448	.3725
29.95	.182	.3952

*in radians, gauss, centim.

Volkmann, 1882

d	σ	d	σ
15°			
0.9991	73.3	1.1798	76.0
1.0400	74.3	.2830	77.5
1.1039	75.0	.3981	80.3

Sentis, 1897

mol%	t	σ
4	16.55	78.6
3	17.6	76.8
2	17.2	75.6
1	16.45	74.6
0	13.5	74.0
0	25.1	72.3

Pann, 1901

%	σ 10°	30°
0	75.37	72.32
10.00	76.68	-
16.17	77.73	74.62
27.63	81.02	77.91
45.9	-	84.23

Brümmer, 1902

%	σ
15°	
0	74.92
2.937	76.02
5.752	77.28
8.487	77.68
11.221	83.05
13.858	87.76

Forch, 1905

M	σ
16°	
0.730	77.41
1.092	77.80
1.734	78.32
2.080	78.72
2.601	79.30
2.913	79.69
3.641	80.63

Stocker, 1920

%	σ
18°	
0	72.55
7.49	73.62
11.11	74.05
15.41	74.89
22.36	76.47

Optical and electrical properties .

Walter, 1889

%	n_D
15°	
0	1.3334
1.88	.3395
12.5	.3596
22.4	.3797
32.6	.4025

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
0°			
0.051	1.32711	1.015	1.35151
.102	.32814	.421	.36081
.203	.33110	.624	.36543
.406	.33618	.827	.36981
.609	.34149	2.032	.37455

Chéneveau, 1907

%	n C	D	Tl	F	G
0	1.33148	1.33330	-	1.33741	1.34062
22.46	1.37554	1.37770	1.37964	1.38225	1.38582

Geffcken, 1929

m	n_{He}	m	n_{He}
25°			
0	1.33259	1.486	1.36945
0.4989	.34600	1.5046	.369875
.9901	.35810	2.0200	.38073
.9901	.35811	2.0214	.38077
1.002	.35842		

Okazaki, 1935

%	Verdet's constant.10^5 (3441 Å)
28°	
6.99	4504
12.62	4595
17.30	4679
22.52	4721
26.97	4805
29.88	4847

Da Silveira, 1939

Raman effect (see author) .

Beetz, 1862

%	κ 28°	κ 50°	%	κ 28°	κ 50°
7.18	253.8	426.0	23.15	492.4	853.2
9.35	304.5	516.8	23.67	493.2	865.8
11.98	363.2	604.7	24.27	492.4	879.0
15.60	416.8	720.3	25.27	491.7	874.2
17.75	473.0	797.9	26.81	482.7	881.4
19.60	478.5	820.5	26.92	482.0	873.0
20.42	481.3	832.0	28.96	465.9	876.3
21.66	488.4	846.2	32.09	426.9	823.9
21.83	493.0	855.3	32.49	420.5	799.8
22.14	493.3	863.5	33.71	403.8	777.8
22.44	491.7	859.5	35.04	380.8	713.5
22.59	491.9	863.6	37.81	339.0	-

t	κ	t	κ
	7.18 %		
9.4	195.8	34.8	342.8
12.0	210.6	41.1	373.7
13.0	216.1	45.0	399.9
15.1	226.8	45.7	401.0
19.9	253.2	54.1	448.9
26.5	291.7	66.1	503.2
28.0	303.1	78.6	551.4
	9.34 %		
11.5	252.5	43.5	469.6
15.3	275.6	48.1	502.1
18.5	294.1	49.7	515.0
22.8	326.2	59.1	573.5
28.0	362.7	70.5	632.8
35.2	412.8	81.0	673.9
	11.88%		
15.7	303.2	41.9	506.4
23.5	357.6	53.3	594.6
27.8	394.8	64.1	667.2
30.4	414.2	73.1	726.8
36.6	461.1	80.8	767.5
	15.60 %		
15.3	353.9	53.7	710.2
20.3	399.6	56.5	743.6
24.5	437.7	69.1	848.8
29.5	487.2	79.5	917.0
37.5	562.4	85.0	954.3
45.7	639.7		
	17.75 %		
10.4	351.4	41.8	664.0
13.2	378.9	50.5	756.0
18.1	426.6	64.9	894.5
22.8	471.9	71.8	951.2
32.2	532.3		
	19.61 %		
13.5	390.2	37.4	646.4
15.9	408.1	51.3	784.6
21.3	463.6	63.3	901.7
25.8	510.7	78.0	1033.6
29.4	543.3		
	20.41 %		
13.9	397.7	43.0	705.1
17.8	429.4	51.3	796.4
21.4	467.1	56.1	845.5
27.3	526.8	75.2	1034.6
33.4	598.3		
	21.66 %		
11.0	364.1	53.3	834.1
16.4	420.1	71.1	1024.1
26.0	519.7	76.3	1070.8
	21.83 %		
8.7	341.0	48.7	791.0
11.6	376.4	60.3	912.2
23.5	500.3	70.6	1019.1
	22.14 %		
13.6	424.6	49.9	862.2
14.5	433.5	58.9	975.7
23.5	531.9	59.0	976.5
26.2	561.1	68.0	1080.2
29.0	597.4	74.5	1146.3
40.7	745.2	77.5	1173.3
43.5	778.6	78.0	1180.3
	22.44 %		
11.6	373.4	43.6	730.4
14.2	395.1	51.8	827.6
16.8	426.1	60.0	921.0
20.6	469.4	73.88	1064.8
26.7	585.5	75.5	1118.9
35.0	629.8		
	22.59 %		
10.3	383.7	35.0	677.6
14.5	428.0	45.2	806.4
18.1	469.3	60.7	991.2
25.3	556.5	79.4	1190.6
	23.15 %		
9.3	368.2	31.9	626.6
14.3	425.9	46.2	806.7
24.9	549.5	71.7	1118.4
	23.66 %		
9.5	376.6	34.4	680.4
11.8	412.0	47.2	850.1
14.5	438.3	62.9	1044.5
20.0	493.2	78.6	1215.2
26.2	578.6		
	24.28 %		
10.7	383.6	31.9	641.6
12.9	406.6	45.0	815.2
14.2	424.5	58.5	987.6
22.7	525.5	78.1	1217.2
	25.28 %		
12.0	391.8	51.1	889.5
18.5	474.4	59.7	1062.2
27.5	583.6	67.1	1094.5
39.9	643.3	71.9	1154.6
40.3	752.2	80.5	1275.9
42.4	766.8		
	26.81 %		
11.7	383.8	44.2	807.4
16.6	441.4	59.6	1003.9
28.0	580.1	73.4	1190.7
32.0	632.7		
	26.92 %		
9.5	352.4	42.2	768.1
13.4	401.9	62.5	1040.6
27.3	570.8		
	28.97 %		
9.3	340.3	46.6	831.6
18.2	444.5	54.6	936.7
23.6	509.0	59.6	1014.5
30.0	589.8	66.7	1097.7
35.8	670.1	83.5	1301.9
42.8	763.6		
	32.09 %		
20.0	426.9	33.3	608.8

	32.50 %		
10.8	321.0	50.8	834.7
17.7	394.2	55.1	894.1
18.6	404.6	69.9	1092.9
21.8	440.9	73.7	1157.0
30.2	547.0	80.5	1254.6
39.5	671.1	82.1	1268.0
44.0	731.9		
	33.71 %		
13.4	322.3	49.3	789.3
13.5	323.1	53.6	853.7
14.1	329.6	57.5	908.3
14.7	337.4	64.1	1009.2
18.0	381.0	66.0	1041.6
20.0	403.8	69.0	1077.5
21.5	420.4	69.5	1088.0
24.7	449.3	73.3	1145.8
29.5	521.3	74.5	1159.3
33.2	560.4	82.5	1253.5
46.0	751.2		
	35.04 %		
0	196.3	20.3	384.3
10.2	278.4	25.8	445.6
11.1	281.1	38.1	613.1
13.5	301.5	46.4	725.6
14.5	312.2	59.5	916.0
15.7	322.6	73.0	1113.4
17.5	352.1	79.2	1199.6
	37.81 %		
10.6	239.1	46.3	660.3
13.8	270.8	56.6	806.7
18.6	318.7	64.5	926.9
23.3	366.8	73.1	1052.4
27.5	412.8	79.1	1136.3
30.5	452.2	82.0	1178.7
37.2	542.0		

Ewing and Mc Gregor, 1872-73

%	κ	%	κ
	10°		
		22.46	330.1
1.37	5.144	23.70	332.4
1.81	66.97	23.78	330.1
2.67	84.66	24.41	327.8
5.10	147.5	26.43	322.8
7.02	185.2	28.07	303.5
9.36	223.4	31.19	293.1
14.04	279.3	31.84	281.8
16.04	293.1	32.04	279.3
18.71	310.4		
21.38	322.8		

Freund, 1879

%	κ 0°	κ 20°	%	κ 0°	κ 20°
4.977	117.64	192.79	19.69	271.60	480.65
9.61	186.00	315.04	22.16	283.51	498.82
14.44	238.95	410.78	27.01	261.30	483.66

Kohlrausch, 1879

%	κ	$\tau . 10^4$
	18°	
5	189	226
10	320	224
15	413	229
20	467	242
25	478	259
30	442	274

Fink, 1885

P	κ 0°	κ 18°	κ 0°	κ 18°
	0.96%		9.67%	
1	38.9	60.9	221.6	360.1
109	40.1	62.3	227.9	368.2
200	41.2	63.5	232.8	374.4
300	42.2	64.9	239.5	381.5
400	43.2	66.2	243.6	388.4
500	44.1	67.3	247.9	394.7
	23.50%		29.64%	
1	336.3	557	288.1	522
109	341.5	564	290.3	527
200	345.8	568	292.5	530
300	350.2	574	294.3	532
400	353.8	579	296.3	535
500	357.3	584	297.7	538

Heim, 1886

t	R	t	R
	39.32 %		
60	208.3	30	429.2
58	216.4	28	457.8
56	223.9	26	485.0
54	232.9	24	516.4
52	244.0	22	552.2
50	255.2	20	590.8
48	266.1	17	656.3
46	279.0	16	679.7
44	292.2	14	735.7
42	306.5	12	796.3
40	322.0	10	858.7
38	339.1	8	938.5
36	356.7	6	1005.1
34	378.3	5	1049.7
32	398.9		
	32.27 %		
45	215.7	8	565.7
40	237.9	7	586.9
35	263.6	6	606.9
30	296.0	4	652.9
25	335.7	1	716.3
20	383.8	0	744.7
15	452.4	-2	800.0
12	496.6	-4	862.5
10	530.3	-6	941.8
9	548.3	-7	984.1

R = relative resistance in arbitrary units .

Jones, 1904; Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.051	52.28	1.015	25.95
0.102	45.84	1.421	18.54
0.203	39.10	1.624	16.24
0.406	32.31	1.827	14.46
0.609	27.90	2.032	12.81

Johnston, 1907

N	λ	N	λ
	99.4°		
0.002	180.5	1	46.0
0.02	100.0	2	27.1
0.2	65.7	4	21.8
0.5	54.3		

Heat constants .

Lundqvist, 1869

%	t	U	heat conductivity
0	40.8	1.0	0.0933
20.82	40.3	0.831	0.0986
21.38	48.0	0.801	0.0942
30.07	45.2	0.770	0.0949

Gray, 1879-80

d	U
at room temp.	
1.327	0.7526
.258	.7755
.161	.8606
.075	.9230

Teudt, 1900

t	U	t	U	t	U
33.20%		16.59%		0%	
31.8	0.6851	32.4	0.8141	34.1	0.9962
34.8	.6998	33.9	.8009	38.1	1.0076
37.9	.6991	34.6	.8176	42.2	.0082
41.3	.7148	38.7	.8096	46.1	.0170
42.9	.7157	41.4	.8103	50.1	.0230
46.1	.7172	43.3	.8175		
47.9	.7252	47.8	.8309		
50.4	.7302	50.9	.8321		
54.1	.7380	54.2	.8283		

Cohen, Helderman and Moesveld, 1924

%	U	%	U
	19°		
0	0.99556	17.14	0.8113
2.15	.9724	25.66	.7354
4.30	.9474	34.21	.6688
8.60	.8989		

Weber, 1880

d	heat conductivity
15°	
0.999	0.0745
1.133	.0711
.271	.0698
.361	.0691

Jäger, 1891

%	relative heat conductivity
0	100
16	95.3
32	91.5

Water + α-Picoline . Zinc chloride ($ZnC_6H_7NCl_2$)

Flaschner, 1909

%	f.t.	sat.t.	%	sat.t.
95.0	95	84.0	40.4	164.6
92.9	91	89.5	30.6	164.6
80.0	-	142.2	21.8	161.7
70.5	-	155.0	16.1	157.5
61.6	-	160.0	12.5	144.0
49.5	91	163.0		

Water + Zinc ammonium sulfate ($ZnH_8N_2O_8S_2$)

Tobler, 1855

%	f.t.	%	f.t.	
4.97	0	10.35	30	
5.91	10	13.03	45	
6.64	13	16.74	60	
8.12	15	20.05	75	(6+1)
8.18	20	23.10	85	

Water + Zinc sodium sulfate ($ZnNa_2O_8S_2$)

Koppel, 1905

%	f.t.
33.21	25 4 aq.
33.30	30
33.28	35
33.47	40

Water + Zinc potassium sulfate ($ZnK_2S_2O_8$)

Tobler, 1855

%	f.t.	%	f.t.
6.72	0	25.61	45
11.92	10	26.52	50
13.89	15	30.51	58
16.91	25	33.92	65
21.57	36 (6+1)	35.39	70

Water + Zinc fluosilicate ($ZnSiF_6$)

Jatlov and Pinaevskaya, 1938

%	f.t.	%	f.t.
0	0	33.73	0 (6+1)
5.0	- 0.8	35.16	20
10.0	- 1.4	37.02	40
15.0	- 3.2	38.30	50
20.0	- 4.8	38.49	60
25.0	- 7.4	40.95	80
30.0	-11.8	42.18	100
32.0	-14.6 E		

Simpson and Glocker, 1953

%	f.t.	%	f.t.
33.67	1.7	38.61	43.3
34.69	10.0	39.43	51.7
35.86	18.3	40.22	60.0
36.88	26.7	40.94	68.3
37.77	35.0		

Water + Zinc formate ($ZnC_2H_2O_4$)

Ashton, Houston and Saylor, 1933

%	f.t.	%	f.t.
3.7	0	11.8	60
4.3	10	15.5	70
5.2	20	21.2	80
6.1	30	28.8	90
7.4	40	38	100
9.2	50		

Water + Zinc acetate ($ZnC_4H_6O_4$)

Favre and Valson, 1874

m	d
23.5°	
0	0.997
0.5	1.057
1.0	1.106
1.5	1.148

de Garcia, 1920

N	d	n_D
22°		
0.062	1.002816	1.3331
0.125	.006448	.3345
0.25	.014033	.3361
0.5	.028082	.3395
1	.054782	.3454
2	.107078	.3569

Water + Zinc benzene sulfonate ($ZnC_{12}H_{10}O_6S_2$)

Ephraim and Seger, 1925

c	f.t.	c	f.t.
14.769	18	32.018	64.5
19.107	34	40.085	80.5
24.705	49.5	42.014	82

Water + Cadmium chloride ($CdCl_2$)

Heterogeneous equilibria

Tammann, 1885

%	p	%	p
	100°		
11.53	746.5	36.11	701.9
19.86	734.2	42.43	681.7
23.55	729.7	49.62	651.8
26.48	724.0	50.62	646.7
32.72	712.7		

Lescoeur, 1894

t	p dissoc. (1+1)	p sat.sol.
10	-	7.4
20	-	12.2
60	-	118
70	-	182
77	-	258
80	50	-
90	84	-
100	151	-
110	232	-

Dieterici, 1923

m	p	m	p
	0°		
0	4.379	1.999	4.393
1.000	.485	2.999	.301
1.491	.441	3.995	.204

Robinson, 1940

Isopiestic solutions

m_1	m_2	m_1	m_2
	25°		
0.1060	0.1054	0.1085	0.1092
.2002	.1870	.2055	.1925
.4285	.3680	.4738	.4021
.7262	.5836	.8772	.6815
1.218	.9120	1.408	1.033
1.818	1.308	1.921	.371
2.201	1.564	2.427	.723
3.273	2.334	3.327	2.383
4.101	3.009	4.298	3.162
5.522	4.241	5.847	4.541
0.1184	0.1180	0.1596	0.1531
.2908	.2597	.3487	.3062
.5778	.4758	.6643	.5366
.9078	.7099	.9990	.7642
1.452	1.065	1.772	1.282
2.120	.509	2.182	.551
2.490	.771	2.710	.925
3.413	2.449	3.754	2.714
4.762	3.554	5.331	4.068
5.993	4.688		

1 - cadmium chloride

2 - potassium chloride

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.622	1.4	0.440
0.2	0.571	1.6	0.435
0.3	0.542	1.8	0.431
0.4	0.522	2.0	0.428
0.5	0.506	2.5	0.430
0.6	0.492	3.0	0.434
0.7	0.481	3.5	0.442
0.8	0.473	4.0	0.454
0.9	0.465	4.5	0.466
1.0	0.458	5.0	0.482
1.2	0.448	6.0	0.514

Johnston, 1906

N	b.t.	N	b.t.
0.330	100.129	4.390	101.324
1.512	100.484	4.936	101.504
2.446	100.754	5.892	101.864
3.182	101.022	6.550	102.110
3.884	101.154		

Etard, 1894

%	f.t.	%	f.t.
43.5	-7	63.0	120
47.6	+1	64.8	150
49.7	6	68.2	165
51.3	7	68.4	170
51.6	10	70.1	180
52.7	19	71.9	190
52.9	25	72.0	200
57.9	61	76.0	235
58.8	82	77.7	270

Dietz, 1899

%	f.t.	%	f.t.
43.58	-9 (4+1)	57.91	36 (1+1)
49.39	0	57.47	10
55.58	+10	57.35	20
59.12	15	57.51	40
44.35	-10 (5+1)	57.77	60
47.37	0	58.41	80
52.53	+18	59.52	100
56.27	30		

Jones; Jones and Chambers, 1900 and 1904

M	f.t.	M	f.t.
0.214	-0.727	0.643	-1.832
.322	-1.022	0.858	-2.329
.429	-1.298	1.072	-2.947

Sudhans, 1914

mol%	f.t.	
10.86	19.3	(5+2)
11.31	29.7	
11.62	40.1	(1+1)
11.63	54.5	

Frederick and Getman, 1929

m	f.t.	m	f.t.
0.0156	-0.0743	0.2023	-0.6733
.0198	.0871	.2288	.7516
.0447	.1822	.3072	.9517
.0611	.2377	.3651	-1.1161
.0801	.3000	.5476	.5549
.1155	.4155	.6577	.8194
.1201	.4287	.7709	-2.0849
.1672	.5764		

Hering, 1936

%	mol%	f.t.	
43.0	6.90	-10.2	E
44.7	7.35	-7.2	(4+1)
47.1	8.05	-3.3	
45.75	7.65	-5.6	tr.t.
46.3	7.81	-3.3	(5+2)
47.3	8.10	0	
50.75	9.20	+12.5	
54.65	10.59	25	
56.95	11.51	32.5	
57.3	11.65	33.5	
57.4	11.69	33.8	
57.4	11.70	37.5	(1+1)
57.5	11.73	45	
57.7	11.81	60	
58.35	12.11	80	
59.55	12.63	100	

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
62.6	141	80.0	261
66.4	168	84.5	290
68.3	182	84.6	330
70.5	196	87.5	383
72.7	207	91.4	433
75.6	220	94.6	481
76.6	227	100	564

Benrath, 1941

%	f.t.	%	f.t.
61.0	124	72	204
63.5	154	73	208
64.5	160	74	214
66.0	166	75.0	218
67.0	169	77.5	233
68.0	173	79.0	252
69	174	82.5	303
70	188	85.5	342
71	198	100.0	564

E: 43.4% -11.5°

tr.t.: 46.2% -5° (4+1) - (5+2)

57.4% +34 (5+2) - (1+1)

69.0 174 (1+1) - anh.

Properties of phases

Kremers, 1858

%	d	%	d
	19.5°		
0	0.998	35	1.394
5	1.043	40	.470
10	.087	45	.558
15	.138	50	.653
20	.193	55	.762
25	.254	60	.887
30	.319		

t	d				
19.5	1.1383	1.2669	1.4017	1.4975	1.7237
0	1.1435	1.2752	1.4128	1.5103	1.7397
40	.1294	.2554	.3873	.4815	.7045
60	.1181	.2414	.3712	.4638	.6838
80	.1045	.2256	.3532	.4444	.6616
100	.0890	.2083	.3340	.4239	.6385

Grotrian, 1883

%	d	%	d
	18°		
1	1.0063	30	1.3305
5	.0436	35	.4075
10	.0919	40	.4878
15	.1443	45	.5775
20	.2007	50	.6799
25	.2620		

Jahn, 1891

c		d
20°		
0		0.9982
19.542		1.1530
29.555	(sic)	.1730
50.066		.2753
37.647		.3177

De Muynck, 1894

%	d	%	d
	18°		
0	0.999	29.977	1.330
14.761	1.142	41.547	.515
21.431	1.210	57.524	.852

Oppenheimer, 1898

%	d	%	d
	20°		
0	0.9982	25.90	1.2435
8.84	1.0801	34.22	.3941
14.88	.1404	44.42	.5645
19.91	.1977		

Biron, 1907

%	d	%	d
	20°		
0	0.9982	28.61	1.3090
0.0607	0.9988	32.52	.3659
0.6045	1.0360	40.65	.4974
5.776	.0504	40.72	.4965
8.046	.0720	45.32	.5828
16.65	.1613	46.29	.6011
23.19	.2380	51.20	.7021
24.24	.2515	51.21	.7020

Getman and Gilroy, 1912

%	d	%	d
	25°		
0	0.9971	18.690	1.1803
1.599	1.0115	21.201	.2165
4.472	.0368	26.076	.2700
8.620	.0695	31.090	.3401
11.600	.1068	41.540	.5153

Rabinowitsch, 1921

%	d	%	d
	25°		
0.95	1.0061	30.63	1.3354
5.19	.0436	35.61	.4083
10.17	.0932	40.79	.4957
15.53	.1472	44.69	.5676
20.61	.2046	48.67	.6457
25.58	.2668	51.62	.7099

Heydweiller, 1921

N	d
18°	
0.5	1.03954
1	.0783
2	.1550
3	.2301
4	.3043

Goard, 1925

M	d
20°	
0.471	1.068
0.942	.136
2.36	.344
3.49	.543
4.71	.672

Spacu and Popper, 1934

%	d	%	d
20°			
0	0.99823	36.664	1.42874
18.333	1.179845	36.750	.43030
27.486	.293414	49.797	.67179

Lüdemann, 1935

N	m	d	N	m	d
25°					
0	0	0.99707	3.6649	1.9488	1.27670
1.0913	0.5555	1.18222	4.3473	2.3405	-
.5300	0.7842	.11585	6.5985	3.7173	1.45236
.9924	1.0288	.15100	8.4761	4.9793	.62808
2.6875	1.4039	.20350	9.9278	6.0485	.73186

Okazaki, 1935

%	d	%	d
28°			
7.06	1.0599	30.43	1.3289
13.77	.1270	36.61	.4220
19.31	.1877	44.73	.5634
25.53	.2632		

Hering, 1936

%	mol%	t	d
47.3	8.10	0	1.635
50.75	9.20	+12.5	.699
54.65	10.59	25.0	.778
56.95	11.51	32.5	.829
57.3	11.65	33.5	.836
57.4	11.69	33.8	.837
57.4	11.70	37.5	.836
57.5	11.73	45	.830
57.7	11.81	60	.820
58.35	12.11	80	.816
59.55	12.63	100	.821

Wagner, 1883

%	η (water°=1)			
	15°	25°	35°	45°
11.09	0.7746	0.6053	0.4910	0.4073
16.30	0.8894	.7047	.5755	.4721
24.786	1.0402	.8037	.6458	.5358

Rabinowitsch, 1921

%	η(water=1)	%	η(water=1)
25°			
0.95	1.022	30.63	1.99
5.19	.086	35.61	2.45
10.17	.182	40.79	3.26
15.53	.304	44.69	4.22
20.61	.460	48.67	5.88
25.58	.680	51.62	7.86

Goard, 1925

M	σ
20°	
0.471	74.04
0.942	74.67
2.36	76.80
3.49	79.20
4.71	82.10

Kremers, 1859

$d^{19.5}$	n_D
15°	
0.9983	1.3334
1.2172	.3699
1.4151	.4026
1.7074	.4504

Jahn, 1891

c	n H_α	n D	n H_β
	20°		
0	1.3315	1.3332	1.3375
19.542	1.3577	1.3599	1.3646
29.555	1.3629	1.3650	1.3700
50.066 (?)	1.3799	1.3828	1.3876
37.647	1.3884	1.3908	1.3964

De Muynck, 1894

%	n_D	%	n_D
	15°		
0	1.33337	29.977	1.38938
14.761	.35835	41.547	.41950
21.431	.37127	57.524	.47314

Getman and Gilroy, 1912

%	n C	n D	n F
		25°	
0	-	1.33250	-
1.599	1.33325	.33517	1.33935
4.472	.33755	.33929	.34361
8.620	.34292	.34499	.34935
11.600	.34858	.35088	.35522
18.690	.36110	.36316	.36811
21.201	.36615	.36833	.37326
26.076	.37631	.37850	.38380
31.090	.38681	.38925	.39473
41.540	.41436	.41710	.42333

Spacu and Popper, 1934

%	n_{He}	%	n_{He}
	20°		
0	1.3324865	36.664	1.403790
18.333	.363114	36.750	.404044
27.486	.381758	49.797	.442730

Lüdemann, 1935

N	m	n_D
	25°	
0	0	1.33254
1.0913	0.5555	.34715
1.5300	0.7842	.35281
1.9924	1.0288	.35868
2.6875	1.4039	.36740
3.6649	1.9488	.37549
4.3473	2.3405	.38781
6.5985	3.7173	.41468
8.4761	4.9793	.43654
9.9278	6.0485	.45320

Jahn, 1891

%	(α)magn.
20°	
0	1
19.542	1.1637
29.555	1.1146
50.066 (sic)	1.1494
37.647	1.1862

Oppenheimer, 1898

%	rotatory magnetic specific polarization
20°	
8.84	1.14
14.88	.14
19.91	.12
25.90	.18
34.22	.14
44.42	.15

Mc Clung and Mc Intosh, 1902

d	% X-Ray absorption
room temperature	
1.310	99.9
.157	99.3
.077	96.6
.038	89.8
.018	77.6
.009	72.9
.000	58.9

Okazaki, 1935

%	Verdet's constant . 10^5 3441 Å
28°	
7.06	4798
13.77	5237
19.31	5618
25.53	6059
30.43	6447
36.61	7057
44.73	7943

Grotrian, 1883

%	κ	τ.10
	18°	
1	54.3	222
5	165	218
10	238	217
15	279	218
20	295	228
25	294	239
30	279	252
35	252	269
40	218	290
45	175	319
50	135	353

Biron, 1907

%	κ
20°	
0.0607	5.95
0.6045	38.0
5.776	178.7
24.24	251.4
40.65	167.0
45.32	161.2

Johnston, 1907

N	λ	N	λ
	99.4°		
0.001	305.3	1	48.0
.10	113.7	2	34.5
.25	84.4	4	20.3
.50	63.5	8	10.0

Rabinowitsch, 1921

%	κ	%	κ
	25°		
0.95	59	30.63	262
5.19	179	35.61	230
10.17	248	40.79	191
15.53	283	44.69	159
20.61	292	48.67	127
25.58	284	51.62	104

Heydweiller, 1921

N	λ
18°	
0.5	28.18
1	19.96
2	12.24
3	8.38
4	5.79

Lilich and Varshavski, 1956

M	pH	M	pH
0.193	6.03	1.12	4.85
.242	5.90	1.70	4.44
.386	5.15	2.27	4.17
.738	5.20	3.90	3.58

Godlewski, 1902

%	e	%	e
	25°		
0.009	0.0040	8.38	0.1654
.041	.0055	13.04	.1758
.055	.0070	16.67	.1806
.092	.0964	21.73	.1851
.36	.1144	28.57	.1898
.91	.1239	33.33	.1929
1.48	.1330	37.50	.1954
2.44	.1397	41.95	.1990
3.64	.1432	44.44	.2012
4.31	.1479	47.37	.2042
5.21	.1528	49.35	.2065
6.10	.1579	51.22	.2084
6.98	.1620	53.10	.2101
7.84			

Helmreich, 1904

%	U		
	30-0°	50-0°	70-0°
0	0.98555	0.98514	0.98966
17.12	.83149	.80956	.81409
32.15	.70187	.69174	.69419
47.65	.56910	.56367	.56577

Water + Cadmium bromide ($CdBr_2$)

Tammann, 1885

%	p
100°	
18.92	745.0
30.91	730.6
42.32	710.7
44.44	705.8
51.23	683.6

Ishikawa and Ueda, 1933

t	p	t	p
	sat.sol.		
(4+1)		anh.	
20	15.90	37.5	40.43
22.5	18.41	40	44.26
25	21.15	42.5	52.68
27.5	24.19		
30	27.59		
32.5	31.50		
35	35.55		

Lescoeur, 1894

t	p	t	p
	sat.sol.		
10	6	60	122
20	10	70	188
30	16	80	279
40	40		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.592	1.2	0.439
0.2	0.533	1.4	0.444
0.3	0.502	1.6	0.449
0.4	0.480	1.8	0.455
0.5	0.466	2.0	0.462
0.6	0.455	2.5	0.483
0.7	0.449	3.0	0.504
0.8	0.445	3.5	0.527
0.9	0.441	4.0	0.548
1.0	0.439		

Robinson, 1940

Isopiestic solutions

m_1	m_2	m_1	m_2
	25°		
0.1065	0.1017	0.1261	0.1187
.2965	.2459	.3897	.3112
.8097	.6037	.8924	.6615
1.437	1.070	1.558	1.168
2.192	1.709	2.285	1.788
2.792	2.263	2.984	2.446
3.639	3.100	3.758	3.224
4.064	3.536		
0.1457	0.1338	0.2205	0.1903
.4563	.3583	0.6634	.4975
.9710	.7122	1.160	.8504
1.884	1.433	2.097	1.617
2.306	1.811	2.515	2.004
3.261	2.717	3.418	2.883
3.844	3.314	5.062	3.530

1 - Cadmium bromide

2 - Potassium chloride

Etard, 1894

%	f.t.	%	f.t.
32.0	-4	60.0	48
34.7	-1	61.2	71
36.3	+1	61.8	104
36.0	2	63.7	155
41.9	9	65.2	170
46.0	14	69.9	215
52.6	25	70.1	232
59.6	35	71.5	245

Dietz, 1899

%	f.t.	%	f.t.
	(4+1)		
37.92	0	56.90	30
48.90	18	61.84	38
	(1+1)		
60.29	35	61.10	60
60.65	40	61.29	80
60.75	45	61.63	100

Jones, 1904 and Jones and Chambers, 1900 and 1904

M	f.t.	M	f.t.
0.22	-0.652	0.66	-1.738
0.44	-1.213	0.88	-2.277

Getman, 1929

m	f.t.	m	f.t.
0.0165	-0.0683	0.3354	-0.9358
.2043	.0951	.4303	-1.1149
.0366	.1367	.4636	.1850
.0680	.2377	.5112	.2944
.0893	.2901	.5232	.3083
.1139	.3683	.5372	.3528
.1994	.5931	.5967	.5004
.2236	.6586	.6360	.5933
.2581	.7368	.6396	.5886

Ishikawa and Ueda, 1933

%	f.t.	%	f.t.
49.64 (4+1)	20	59.66	35
51.26	22.5	60.32 anhydre	35
52.88	25	60.33	36
54.53	27.5	60.39	37
56.21	30	60.41	38
58.61	33.5		

Hering, 1936

%	mol%	f.t.	
32.90	3.14	-4.4	E
36.00	3.59	0	(4+1)
52.9	6.91	25	
57.85	8.33	32.5	
60.3	9.13	36.0	tr.t.
60.3	9.14	38	anh.
60.35	9.15	45	
60.45	9.19	60	
60.8	9.31	75	
61.65	9.62	100	

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
64.6	153	74.8	266
65.7	164	79.2	312
66.5	174	89.6	419
71.8	237	100	567

Benrath, 1941

%	f.t.	%	f.t.
67.2	185	75.0	269
68.0	196	77.5	290
69.1	210	82.9	350
70.1	218	100.0	567
72.5	248		

E : 33.15% -4.4° tr.t. 60.33% +36°

Kremers, 1856 and 1859

%	d	%	d
	19.5°		
0	0.998	30	1.324
5	1.041	35	.398
10	.088	40	.479
15	.139	45	.575
20	.197	50	.677
25	.258		

t	d		
19.5	1.2549	1.4406	1.6386
0	1.2621	1.4518	-
40	.2441	.4262	1.6204
60	.2310	.4100	.6011
80	.2158	.3921	.5803
100	.1990	.3728	.5582

Grotrian, 1883

%	d	%	d
	18-22°		
1	1.0072	25	1.2605
5	.0431	30	.4052
10	.0907	40	.4915
15	.1432	43	.5467
20	.1991		

Jahn, 1891

c	d
20°	
0	0.982
19.181	1.1607
39.350	1.3289

De Muynck, 1894

%	d	%	d
	18°		
0	0.999	11.983	1.112
1.927	1.017	20.552	.209
3.734	.030	23.973	.252
6.432	.106	33.289	.384

Le Blanc and Rohland, 1896

%	d
20°	
0	0.9982
18.06	1.1358
21.39	1.1645

Forchheimer, 1899

%	d	%	d
	20°		
0	0.9982	37.53	1.4469
12.46	1.1211	46.574	1.6198
22.53	1.2293		

Getman and Gilroy, 1912

%	d	%	d
	25°		
0	0.9971	15.66	1.1509
1.05	1.0059	21.13	.2133
2.76	.0212	27.51	.2991
5.41	.0470	40.74	.4918
11.03	.1016		

Hering, 1936

%	mol %	t	d
	sat.sol.		
32.90	3.14	- 4.4 E	-
36.00	3.59	0 (4+1)	1.441
52.90	6.91	+25	1.775
57.85	8.33	32.5	1.904
60.3	9.13	36 tr.t.	1.974
60.3	9.14	38 anh.	1.971
60.35	9.15	45	-
60.45	9.19	60	1.949
60.8	9.31	75	1.938
61.65	9.62	100	1.928

Rabonowitsch, 1936

%	d	η (water=1)
	25°	
5.25	1.0447	1.050
:9.21	.0836	1.097
14.20	.1339	1.154
18.19	.1787	1.212
24.71	.2585	1.34
31.30	.3510	1.51
37.83	.4569	1.76
42.23	.5377	2.02
46.63	.6279	2.38
51.26	.7334	2.96
	50°	
6.35	1.044	1.064
12.0	.098	1.129
18.0	.163	1.207
24.1	.235	1.301
30.2	.317	1.44
36.1	.406	1.62
43.4	.538	1.97
50.9	.693	2.57
55.8	.814	3.26
60.8	.958	4.47

Kremers, 1859

$d^{19.5}$	t	n_D
0.9983	16	1.3333
1.3288	14	.3814
1.5924	16	.4199

Jahn, 1891

c	n		
	H_α	D	H_β
		20°	
0	1.3315	1.3332	1.3375
19.181	1.3550	1.3572	1.3621
39.350	1.3793	1.3817	1.3876

De Muynck, 1894

%	n_D	%	n_D
	15°		
0	1.33337	20.552	1.36555
1.927	.33665	23.973	.37180
3.734	.33916	33.289	.39215
6.543	.34309	41.961	.41386
11.983	.35125		

Le Blanc and Rohland, 1896

%	n_D
	20°
0	1.3333
18.06	1.3532
21.39	1.3579

Getman and Gilroy, 1912

%	n		
	C	D	F
		25°	
0	-	1.33250	-
1.05	1.33043	.33330	1.33891
2.76	.33430	.33623	.34038
5.41	.33834	.34001	.34458
11.03	.34599	.34782	.35247
15.66	.35308	.35500	.36008
21.13	.36213	.36425	.36965
27.51	.37381	.37624	.38184
40.74	.40130	.40373	.41063

Jahn, 1891

c	$(\alpha)_{magn.}$
20°	
0	1
19.181	1.3128
39.350	1.2960

Forchheimer, 1899

%	(α) magn.
20°	
12.46	1.335
22.53	.321
37.53	.329
46.574	.336

Grotrian, 1883

%	κ	τ
	18-22°	
1	35.2	232
5	107	226
10	161	232
15	202	236
20	232	239
25	254	247
30	269	258
35	274	270
40	267	281
43	257	288

Johnston, 1907

N	λ
99.4°	
0.005	314.5
0.2	92.7
0.5	62.6
1	46.2
2	34.2
4	23.6
6	16.9

Rabinowitsch, 1936

%	κ	%	κ
	25°		
0	0	31.30	310
5.25	124	37.83	309
9.21	177	42.23	297
14.20	229	46.63	275
18.19	254	51.26	243
24.71	290		
	50°		
0	0	36.1	510
6.35	215	43.4	501
12.0	311	50.9	454
18.0	389	55.8	402
24.1	449	60.8	339
30.2	491		

Lilich and Varshavski, 1956

M	pH	M	pH
0.09	6.30	0.602	5.47
.150	6.22	1.005	5.06
.288	5.97	1.964	4.52
.320	5.94	3.00	3.97
.426	5.77		

Helmreich, 1904

%	U		
	35-0°	50-0°	70-0°
0	0.98555	0.98514	0.98966
16.13	.84123	.83002	.83357
31.88	.69693	.68799	.68968
41.98	.60533	.60066	.60204

Water + Cadmium iodide (CdI_2)

Heterogeneous equilibria .

Tammann, 1885

%	p
100°	
22.69	747.6
30.23	742.5
37.33	733.6
46.65	718.8
53.09	704.1

Robinson and Wilson, 1940

Isopiestic solutions

m_1	m_2	m_1	m_2
25°			
0.08585	0.05736	0.1227	0.08092
.1817	.1155	.2338	.1446
.4188	.2498	.6258	.3811
.8279	.5234	.9578	.6125
1.432	.9891	1.578	1.108
1.675	1.194	1.725	.233
2.143	1.598	2.317	.773
0.1559	0.1021	0.1774	0.1129
.3111	.1862	.4163	.2463
.7164	.4412	.8135	.5113
1.094	.7059	1.147	.7485
.597	1.112	.607	1.110
.842	.346	.955	.439
2.423	.893	2.432	.910

1 - cadmium iodide 2 - potassium iodide

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
25°			
0.1	0.416	0.9	0.388
0.2	0.390	1.0	0.394
0.3	0.371	1.2	0.405
0.4	0.365	1.4	0.419
0.5	0.366	1.6	0.436
0.6	0.370	1.8	0.451
0.7	0.376	2.0	0.463
0.8	0.383	2.5	0.507

Beckmann, 1890

%	b.t.	%	b.t.
4.16	100	11.31	100.181
4.48	100.073	12.51	100.212
7.83	100.121	17.49	100.303
8.76	100.143	19.25	100.353

Johnston, 1906

%	b.t.	%	b.t.
4.296	100.077	26.05	100.556
7.22	100.131	28.72	100.633
11.05	100.191	31.15	100.725
13.65	100.243	33.45	100.824
16.27	100.308	35.78	100.925
20.17	100.381	37.90	101.022
23.24	100.461	42.15	101.099

N	b.t.	N	b.t.
0.445	100.169	1.308	100.436
0.639	100.218	1.563	100.517
1.048	100.286	1.825	100.577
1.210	100.354	2.021	100.682

Rüdorff, 1872

%	f.t.
9.09	-0.50
16.67	-1.10
26.47	-2.20
31.51	-3.05
33.33	-3.40

Dupuy, 1884

%	f.t.
44.50	15
45.10	30
50.07	60
52.45	100

Etard, 1894

%	f.t.	%	f.t.
42.4	-4	55.1	94
43.7	-2	54.7	95
45.2	-10E	62.9	135
44.8	13	63.1	140
46.5	24	68.1	165
47.4	32	70.7	185
49.5	54	73.4	202
50.1	64	73.2	202
52.4	76	81.5	255

Dietz, 1899

%	f.t.
44.39	0
46.02	+18
49.35	50
52.65	75
56.08	100

Chambers and Frazer, 1900 and Jones, 1904

M	f.t.
0.133	-0.314
0.222	-0.479
0.333	-0.710
0.444	-0.997
0.666	-1.564
0.888	-2.227

Cohen and Brunsis, 1918 - 1919

%	f.t.
44.18	0
45.20	12
45.72	18
46.00	21
46.46	26

Cohen, Hetterschy and Moesveld, 1920

%	f.t.	%	f.t.
44.083	0.00	46.793	30.00
45.363	15.00	47.039	32.50
45.829	20.00	47.296	35.00
46.053	22.50	47.573	37.50
46.295	25.00	47.833	40.00
46.540	27.50		

Getman, 1929

mol%	f.t.	mol%	f.t.
0.0104	- 0.0357	0.2548	-0.5206
.0211	.0673	.2702	.5426
.0323	.0953	.2782	.5585
.0468	.1267	.3277	.6595
.0699	.1743	.3377	.6714
.0822	.2010	.3620	.7278
.0874	.2089	.4263	.8606
.1075	.2426	.4808	.9705
.1412	.2991	.5039	1.0180
.1732	.3585	.9847	2.328

Guempel, 1929

%	f.t.	%	f.t.
5.37	-0.2	42.19	-5.45
15.0	0.8	44.08	6.0
20.11	1.3	45.60	6.4
25.50	2.0	46.09	+20.0
33.0	3.4	48.11	40.0
36.80	4.0	55.06	100
37.73	4.2	77.15	240
39.66	4.7	100	385
40.90	-5.0		

Hering, 1936

%	mol %	f.t.
42.05	3.445	- 5.3 E
43.35	3.625	- 3.5 (4+1)
44.4	3.78	- 2.0
43.75	3.675	- 2.9 tr.t.
43.85	3.70	- 1.8 anh.
44.05	3.73	0
45.85	4.0	+20
47.95	4.335	40
50.1	4.71	60
52.6	5.18	80
55.55	5.79	100

Benrath, Gjèdebo and al., 1937

%	f.t.	%	f.t.
59.1	128	85.4	269
66.7	179	86.0	271
73.3	211	90.5	293
75.4	223	94.8	318
79.5	243	100.0	385

Eddy and Menzies, 1940

mol%	f.t.
5.37	76.9
6.18	99.1
7.58	131.2
8.75	155.1

Properties of phases.

Kremers, 1858 and 1860

%	d	%	d
	19.5°		
0	0.908	30	1.317
5	1.042	35	.393
10	.086	40	.474
15	.136	45	.572
20	.192	50	.678
25	.251		

Kremers, 1858 and 1860

t	d	
0	1.31448	-
19.5	.30740	1.58960
40	.29621	.57344
60	.27928	.55541
80	.26291	.53550
100	.20452	.51398

Grotrian, 1883

%	d	%	d
	18°		
1	1.0071	25	1.2550
5	.0425	30	.3228
10	.0883	35	.4000
15	.1392	40	.4816
20	.1943	45	.5741

Jahn, 1891

c	d
	20°
0	0.9982
18.744	1.1520
34.102	.2770
63.526	.5154

De Muynck, 1894

%	d
	15°
0	0.999
9.559	1.086
13.677	1.125
31.123	1.338

Le Blanc and Rohland, 1896

%	d
	20°
0	0.99823
10.97	1.0962
16.53	1.1541

Forchheimer, 1900

%	d
	20°
0	0.9982
7.26	1.0630
14.40	.1355
27.07	.2837
44.53	.5807

Dinkhauser, 1905

%	d	%	d
	18°		
2.162	1.0171	20.342	1.1983
5.514	.0467	25.388	.2592
8.760	.0767	27.204	.2828
17.200	.1627		

Agerer, 1905

%	d
	18°
8.851	1.0757
20.342	.1973
27.204	.2828

Doroszewski and Rakowski, 1909

%	d	%	d
	15°		
0	0.9991	25	1.2557
5	1.0431	30	.3234
10	.0876	35	.3979
15	.1389	40	.4828
20	.1947	45	.5766

Getman and Gilroy, 1912

%	d	%	d
	25°		
0	0.9971	13.491	1.1216
0.834	1.0042	20.133	.1933
3.340	.0249	26.785	.2780
3.880	.0303	39.860	.4801
6.984	.0581		

Cohen, Hetterschij and Moesveld, 1920

%	d	%	d
	30°		
0	0.99567	32.014	1.34529
8.433	1.06814	40.131	.47566
15.998	.14493	46.793	.60269
24.533	.24381		

Cohen, Helderman and Moesveld, 1924

%	d
30°	
8.43	1.06922
17.41	.16063
28.21	.29178
35.41	.39690
44.30	.55289

Hering, 1936

%	t	d
44.05	0	1.567
45.85	20	1.590
47.95	40	1.619
50.1	60	1.646
52.6	80	1.682
55.55	100	1.726

Rabinowitsch, 1936

%	d	%	d
	25°		
		24.97	1.2520
4.88	1.0393	31.49	.3405
9.54	.0817	37.95	.4414
14.13	.1273	41.68	.5062
19.10	.1810	44.43	.5581

Gibson, 1937

%	d	%	d
	25°		
2.838	1.02103	19.993	1.19145
4.990	.03975	24.574	.24524
5.625	.04557	29.523	.31251
7.530	.06270	39.396	.46614
9.847	.08453	40.324	.48269
15.446	.14070		

Guillaume, 1946

%	d
20°	
6.08	1.0531
11.0	.1049
29.3	.3140
42.19	.5206

Gibson, 1937

%	π (1-1000 bars)
25°	
0	39.95
2.838	39.03
4.990	38.83
5.625	38.71
7.530	38.48
15.446	37.47
24.574	36.08
40.324	33.77

Rabinowitsch, 1936

%	η(water=1)	%	η(water=1)
25°			
0	1	24.97	1.268
4.88	1.042	31.49	.39
9.54	.079	37.95	.85
14.13	.122	41.68	.68
19.10	.184	44.43	.79

Kremers, 1859

$d^{19.5}$	n_D
16°	
0.9983	1.3333
1.2912	.3792
1.5956	.4293

Barbier and Roux, 1890

c	dispersive power*
11°	
9.885	0.400
19.77	.453
28.95	.503
40.00	.568

* $\frac{n_a - n_b}{1/a^2 - 1/b^2}$ where a= 4524 Å and b= 6452 Å

Forchheimer, 1900

%	(α) magn.
20°	
7.26	2.086
14.40	.094
27.07	.102
44.53	.159

Jahn, 1891

c	$(\alpha)_{magn.}$
20°	
0	1
18.744	1.9667
34.102	2.0002
63.526	2.0731

Agerer, 1905

%	t	$(\alpha)_{magn.}$
8.851	17.7	1.983
20.342	17.9	2.015
27.204	17.8	2.040

Guillaume, 1946

%	$(\alpha)^*_{magn.}10^6$	5780Å n
20°		
6.08	4.210	1.3429
11.0	4.438	.3515
29.3	5.204	.3859
42.19	5.821	.4202

* in radians, gauss, centim.

Jahn, 1891

c	n H_α	n D	n H_β
20°			
0	1.3315	1.3332	1.3375
18.7	1.3556	1.3580	1.3634
34.102	1.3755	1.3782	1.3847
63.526	1.4142	1.4176	1.4263

De Muynck, 1894

%	n_D	%	n_D
15°			
0	1.33337	13.677	1.35474
3.095	.33822	18.728	.36370
3.253	.33871	24.221	.37449
3.379	.33885	31.123	.38999
9.559	.34801	35.482	.40085
12.723	.35329	39.959	.41332

Le Blanc and Rohland, 1896

%	n_D
20°	
0	1.3333
10.97	.3488
16.53	.3582

Dinkhauser, 1905

%	n_D	%	n_D
	18°		
2.162	1.33652	20.342	1.36560
5.514	.34128	25.388	.37539
8.760	.34624	27.204	.37922
17.200	.35998		

Getman and Gilroy, 1912

%	n C	n D	n F
		25°	
0	-	1.33250	-
0.834	1.33186	.33368	1.33804
3.340	.33548	.33714	.34151
3.880	.33547	.33780	.34193
6.984	.34021	.34228	.34699
13.491	.35049	.35258	.35706
20.133	.36167	.36435	.37007
26.785	.37553	.37839	.38508
39.860	.40798	.41104	.41952

Grotrian, 1883

%	κ	τ
	18-22°	
1	20.9	0.0286
5	50.1	0.0260
10	102.5	0.0248
15	143	0.0241
20	183	0.0240
25	219	0.0241
30	250	0.0244
35	278	0.0248
40	298	0.0253
45	309	0.0259

Johnston, 1905

N	λ	N	λ
	100°		
0.001	351.6	1.0	43.0
0.1	104.3	2.0	34.3
0.2	80.0	3.20	28.5
0.5	53.0		

Rabinowitsch, 1936

%	κ	%	κ
	25°		
0	0	24.97	257
4.88	70	31.49	306
9.54	116	37.95	345
14.13	160	41.68	360
19.10	206	44.43	368

Lilich and Varshavski, 1956

M	pH	M	pH
0.118	6.34	0.515	5.70
0.162	6.19	0.733	5.52
0.253	6.03	1.38	5.09
0.316	5.94	1.933	4.76
0.441	5.82		

Helmreich, 1903

%	U 30-0°	U 50-0°	U 70-0°
0	0.98555	0.98514	0.98966
15.16	.87235	.84841	.83952
29.99	.72725	.71462	.70932
45.70	.57953	.57220	.56936

Timofeev, 1905

%	U
16.75	0.8443
44.85	0.5884

Doroszewski and Rakowski, 1909

%	U	%	U
		15°	
5	0.9402	30	0.7107
10	.8880	35	.6657
15	.8398	40	.6200
20	.8031	45	.5743
25	.7592		

Cohen, Helderman and Moesveld, 1924

%	U
	19°
8.43	0.9250
17.41	.8410
28.21	.7378
35.41	.6647
44.30	.5877

Cohen and Bruins, 1918 - 1919

Q dil. (one mole in a great amount of sat. sol.) at 18° = -1236

Water + Cadmium ammonium chloride (CdH_4NCl_3)

Rimbach, 1897

%	f.t.	%	f.t.
29.94	2.4	43.99	63.8
33.45	16.0	52.23	105.9
38.96	41.2		

Water + Cadmium ammonium bromide ($CdNH_4Br_3$)

Rimbach, 1905

%	f.t.	%	f.t.
53.82	1.0	65.31	52.2
58.01	14.8	75.985	110.1

Water + Cadmium sodium chloride ($CdCl_4Na_2$)

Sudhaus, 1914

mol %	f.t.	mol %	f.t.
3.40	19.3	3.98	40.1
3.66	29.7	4.28	54.5

Water + Cadmium potassium chloride ($KCdCl_3$)

Rimbach, 1897

%	f.t.	%	f.t.
21.87	2.6	40.67	60.6
26.60	15.9	51.67	105.1
35.66	41.5		
		(1+1)	

Sudhaus, 1914

%	f.t.	%	f.t.
27.49	19.3	34.78	40.1
31.50	29.7	38.23	54.5

Water + Cadmium rubidium chloride ($RbCdCl_3$)

Rimbach, 1902

%	f.t.	%	f.t.
12.97	1.2	30.83	57.6
16.80	14.5	46.62	103.9
25.31	41.4		

Water + Cadmium magnesium chloride (Cd_2MgCl_6)

Rimbach, 1897

%	f.t.	%	f.t.
45.61 (12+1)	2.4	58.14	67.2
49.69	20.8	65.48	121.8
53.51	45.5		

$p(2.4° - 121.8°) = 45.98 + 0.16505\ t.$

Water + Cadmium barium chloride ($CdCl_4Ba$)

Rimbach, 1897

%	f.t.	%	f.t.
41.88	22.5	52.28	62.0
44.67	32.9	62.05	97.8
46.81	41.4	64.83	108.3
50.30	53.4	65.31	109.2
	(4+1)		

Water + Dicadmium barium chloride (Cd_2Cl_6Ba)

Rimbach, 1897

%	f.t.	%	f.t.
45.60	22.6	54.47	69.5
49.14	41.3	62.29	107.2
51.04	53.9	62.48	107.2
53.08	62.2		
		(5+1)	

Water + Cadmium sodium bromide ($Cd_3Br_8Na_2$)

Jones and Knight, 1899

M	λ	M	λ
25°			
4.076	132.6	0.0512	548.3
2.046	187.5	.0256	625.3
1.023	248.1	.0128	701.6
0.5115	315.9	.0064	780.8
.2558	387.0	.00256	856.8
.1279	457.7	.00128	910.7
.1023	478.3	.00064	960.4

Water + Cadmium potassium bromide ($CdBr_3K$)

Rimbach, 1905

%	f.t.
53.635	0.4
58.61	15.8
67.875	50.0
78.11	112.5

Water + Cadmium monorubidium bromide ($CdBr_3Rb$)

Rimbach, 1905

%	f.t.
36.925	0.4
41.46	14.5
56.61	49.2
76.68	107.5

P dissoc. (0–107°) = 35.34 + 0.393 t.

Water + Cadmium tetrarubidium bromide ($CdBr_6Rb_4$)

Rimbach, 1905

%	f.t.
49.48	0.5
54.805	13.5
66.645	51.5
80.055	114.5

P dissoc. (0–115°) = 50.88 + 0.2637 t.

Water + Cadmium potassium iodide (CdI_4K_2)

Jones and Caldwell, 1901

c	f.t.	c	f.t.
5.051	-0.064	25.256	-3.177
7.577	-0.960	30.308	-3.878
12.628	-1.557		

Grotrian, 1883

%	d	κ	$\tau.10^4$
18°			
1	1.0065	40.5	235
5	.0384	155	227
10	.0808	292	224
15	.1269	430	218
20	.1770	571	215
25	.2313	723	214
30	.2890	885	211
35	.3557	1049	207
40	.4282	1223	203
45	.5065	1397	198

Jones and Caldwell, 1901

N	λ	N	λ
18°			
0.5	151.6	0.005	313.0
.25	166.0	.0025	344.8
.10	188.0	.0010	378.0
.05	208.0	.0005	401.0
.025	228.8	.00025	408.0
.010	285.4		

Water + Cadmium strontium iodide (CdI_4Sr)

Jones and Caldwell, 1901

M	molecular conductivity	M	molecular conductivity
15°			
0.00020	462.6	0.025	235.4
0.00025	454.5	0.050	199.9
0.00050	426.4	0.100	172.5
0.00100	393.2	0.250	144.4
0.00250	357.6	0.500	121.9
0.00500	324.4	1.000	91.1
0.01000	289.1		

Water + Cadmium chlorate (CdO_6Cl_2)

Tammann, 1885

%	p
100°	
11.80	743.3
22.47	717.9
28.10	700.4
35.38	668.9
42.81	624.7
60.41	449.1

Meusser, 1902

%	f.t.	%	f.t.
-6.5	26.18	0	74.95
-13	52.36	+18	76.36
-20	72.18	49	80.08
-15	72.53	65	82.95

d sat.sol.: 18°= 2.284 (2+1)

Traube, 1895

%	d
15°	
0	0.99913
1.796	1.01361
5.205	.04176
9.819	.08218
15.184	.13116

Rubien, 1911

N	d
18°	
0	0.9986
0.555	1.0598
1.110	.1200
2.220	.2394
4.440	.4727

Heydweiller, 1912

N	d
18°	
0.111	1.01081
0.222	.02311
0.555	.05982
1.110	.1200
2.220	.2394
4.440	.4727

Heydweiller, 1909

N	n_D
18°	
0	1.33327
0.5	.34061
1.0	.34775
2.0	.36193
4.0	.38842

Rubien, 1911

N	n_D
18°	
0	1.33327
0.555	.34140
1.110	.34929
2.220	.36504
4.440	.39484

Heydweiller, 1912

N	$\varkappa$
18°	
0.111	82.5
0.222	153.6
0.555	339.5
1.110	589.4
2.220	924.5
4.440	1097

Water + Cadmium nitrate (CdN_2O_6)

Tammann, 1885

%	p	%	p
	100°		
7.56	749.3	37.65	657.8
14.99	734.9	37.80	656.7
24.35	708.2	43.92	622.2
32.50	681.0	46.44	600.5

Dieterici, 1923

m	p	m	p
	0°		
0	4.579	1.994	4.083
0.501	.484	2.500	3.927
0.998	.368	3.004	3.756
1.496	.233	3.904	3.426

Ewing and Guyer, 1938

t	p sat.sol.	t	p sat.sol.
	(4+1)		
0.00	2.93	57.13	40.14
4.94	4.09	58.00	39.39
9.97	5.56	58.96	32.88
15.00	7.34	58.00	27.91
19.94	9.81	57.75	26.10
24.88	12.76	55.01	18.73
29.87	16.36	54.86	18.56
35.03	20.68	49.99	11.58
40.18	25.77	44.89	7.42
45.18	30.95	44.62	7.42
50.03	36.09	39.97	4.80
55.06	39.97	34.98	3.06
56.13	40.22		
	(2+1)		
30.00	5.26	53.04	11.31
40.03	7.95	54.02	11.48
49.02	10.59	55.03	11.34
49.93	10.81	55.91	11.20
51.04	11.05	56.02	11.00
52.00	11.26	56.60	10.92
52.42	11.33		
	anhydre		
20.01	1.18	55.02	8.35
25.00	1.56	60.00	10.86
30.00	2.13	61.43	11.58
35.00	2.84	61.60	11.70
39.98	3.71	70.13	12.37
50.03	6.45		

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.850	0.9	0.946
0.2	0.852	1.0	0.962
0.3	0.861	1.2	0.995
0.4	0.873	1.4	1.025
0.5	0.888	1.6	1.057
0.6	0.903	1.8	1.085
0.7	0.917	2.0	1.114
0.8	0.931	2.5	1.182

Rüdorff, 1872

%	f.t.
4.47	-0.90
9.83	-2.30
15.83	-4.20
19.61	-5.90
25.47	-8.65

Funk, 1899

%	f.t.		%	f.t.	
30.68	-9.5		52.31	0	(4+1)
35.62	-13		55.90	18	
35.91	-14.5		58.40	30	
37.37	-13	(9+1)	61.42	40	
47.33	-1		76.54	59.5	
52.73	+1				

Jones and Getman, 1904
Jones, 1904 and Jones and Bassett, 1905

N	f.t.	N	f.t.
0.0845	-0.41	1.0146	- 5.73
0.1691	-0.85	1.6910	-10.58
0.3382	-1.73	2.3674	-17.00
0.6764	-3.50		

Vasilyev, 1910

E : 81.37 % 44.5°

(4+1) 59.5°

There is also an (8+1), but (2+1) does not exist

Sieverts and Petzold, 1933

%	f.t.	%	f.t.
24.2	-7.2	73.0	55.0
31.7	-11.5	74.4	58.0
41.4	-20.0	78.3	58.0
36.9	-16.0E	79.9	55.0
38.2	-14.1 (9+1)	82.3	48.7 E
41.6	-10.0	82.5	50.0 (2+1)
45.9	-5.2	82.7	51.0
51.8	+0.3	84.4	55.0
56.1	3.5tr.t.	86.0	56.8 tr.t.
55.1	0.6(4+1)	86.1	61.0 anh.
58.7	15.0	87.0	79.0
61.3	25.0	87.2	100.0
64.3	35.0	88.2	130
70.0	50.0		

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987	26	1.2970
1	1.0093	27	.3125
2	.0199	28	.3260
3	.0305	29	.3405
4	.0411	30	.3548
5	.0514	31	.3710
6	.0604	32	.3872
7	.0694	33	.4034
8	.0784	34	.4196
9	.0874	35	.4354
10	.0964	36	.4553
11	.1072	37	.4753
12	.1180	38	.4953
13	.1288	39	.5153
14	.1395	40	.5353
15	.1501	41	.5572
16	.1625	42	.5792
17	.1749	43	.6012
18	.1872	44	.6292
19	.1996	45	.6454
20	.2118	46	.6680
21	.2260	47	.6907
22	.2402	48	.7134
23	.2543	49	.7361
24	.2685	50	.7586
25	.2825		

Grotrian, 1883

%	d	%	d
	18°		
1	1.0069	30	1.3125
5	.0415	35	.3802
10	.0869	40	.4590
15	.1360	45	.5430
20	.1903	48	.5978
25	.2500		

De Muynck, 1894

%	d	%	d
	18°		
0	0.999	30.876	1.321
8.683	1.074	43.716	.515
14.899	.134	54.027	.711
21.353	.204		

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	0.9963	1.09398
0.1012	1.00854	1.985	.18733
.2017	.01814	3.984	.37032
.4997	.04686		

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
18.465	1.1435
36.93	1.2874

Jones, 1904 and Jones and Getman, 1904

N	d	N	d
	0°		
0.0845	1.017168	1.0146	1.206336
0.1691	.034032	1.6910	.331776
0.3382	.067928	2.7056	.539228
0.6764	.137908		

Jones and Bassett, 1905

N	d	N	d
	0°		
0.0845	1.015340	1.0146	1.189500
0.1691	.034084	1.6910	.289544
0.3382	.066612	2.3674	.423460
0.6764	.124180		

Guillaume, 1946

%	d	%	d
	20°		
13.42	1.1185	31.8	1.3261
20.35	.1898	38.0	.4144
29.8	.3025	41.0	.4594

Ewing and Herty, 1953

%	d 25°	30°	40°	50°	60°
0.0	0.9970	0.9954	0.9919	0.9878	0.9832
19.747	1.1824	1.1798	1.1746	1.1692	1.1637
20.458	.1901	.1875	.1824	.1769	.1712
31.149	.3208	.3177	.3113	.3049	.2983
41.221	.4680	.4646	.4577	.4505	.4432
43.076	.4977	.4937	.4860	.4782	.4705
50.654	.6352	.6308	.6220	.6129	.6036
54.313	.7080	.7033	.6941	.6850	.6756
60.665	.8505	.8454	.8352	.8252	.8151
61.560	.8731	.8681	.8581	.8482	.8381
65.813	.9803	.9753	.9648	.9547	.9445
67.866	2.0360	2.0305	2.0198	2.0094	.9990
72.373	.1698	.1640	.1521	.1397	2.1276
75.486	.2684	.2628	.2517	.2403	.2290
80.274	.4272	.4211	.4090	.3968	.3845
83.903	.5666	.5594	.5468	.5338	.5209
84.384	.5817	.5760	.5641	.5521	.5404
85.064	.6114	.6052	.5929	.5801	.5672
87.395	.7065	.6995	.6855	.6712	.6571

Wagner, 1883

%	η (water°=100) 15°	25°	35°	45°
7.81	61.95	50.13	41.12	34.04
15.71	71.80	58.75	48.76	41.31
22.36	85.11	69.03	57.29	47.53

De Muynck, 1894

%	n_D 15°	20°
0	1.33337	1.33298
8.683	.34518	.34451
14.899	.35386	.35303
21.353	.36323	.36256
30.876	.37904	.37835
43.716	.40453	.40393
54.027	.42920	.42857

Jones, 1904; Jones and Getman, 1904; Jones and Bassett, 1905

N	n_D	N	n_D
	0°		
0.0845	1.32355	1.6910	1.36794
.1691	.33070	2.3674	.37379
.3382	.33560	2.7056	.39985
.6764	.34588	3.0438	.40792
1.0146	.35580	3.3820	.41539

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.9963	1.34689
0.1012	.33470	1.985	.35993
.2017	.33608	3.984	.38487
.4997	.34018		

Heydweiller, 1912

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34694
0.1	.33468	2.0	.36013
0.2	.33507	4.0	.38508
0.5	.34018		

Guillaume, 1946

%	t	$(\alpha)^*_{magn.}10^6$	n
13.42	21.5	3.598	1.3504
20.35	23	.404	.3601
29.80	21	.135	.3752
31.80	20	.082	.3799
38.00	23	2.907	.3900
41.00	21.5	2.824	.3966

*in radians, gauss, centim.

de Mallemann and Guillaume, 1945

$(\alpha)_{magn.}10^5$ ($d^{20°}$= 1.3007)=25.98

Grotrian, 1883

%	κ	$\tau.10^4$	%	κ	$\tau.10^4$
		18-22°			
1	68.5	226	30	947	214
5	286	221	35	938	220
10	507	215	40	895	228
15	680	213	45	814	242
20	817	212	48	747	252
25	909	213			

Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.0845	87.17	1.0146	50.22
.1691	83.55	1.6910	34.27
.3382	73.42	2.3674	23.47
.6764	59.00		

Water + Cadmium perchlorate ($CdCl_2O_8$)

Hering and Leroy, 1939

mol%	f.t.	mol%	f.t.
1.26	-4.4	14.3	129.4
2.24	-10.1	15.8	126.1
3.14	-18.9	17.2	115.2
4.03	-29.7	18.65	100.0
4.27	-33.9	19.9	80.0
4.52	-37.8	21.15	60.0
4.79	-42.5	21.3	58.5 tr.t.
5.07	-48.1	21.3	58.7 (2+1) I
5.37	-55.5	21.7	64.1
5.72	-66.5 E	21.8	65.0
5.81	-61.2(6+1)	21.85	66.0 tr.t.
5.97	-49.5	21.6	60.0 (2+1)II
6.18	-37.3	21.9	67.0
6.26	-34.3	22.5	80.0
6.66	-14.8	23.6	100.0
6.82	-7.2	24.7	115.2
7.17	+8.7	25.6	126.1
7.47	+20.0	25.9	128.7
7.73	30.0	28.55	148.7
8.02	40.0	29.2	153.3
8.33	50.0	33.3	157.9
8.67	60.0	35.8	155.9
9.05	70.0	38.4	150.0
9.46	80.0	40.0	144.4
9.94	90.0	41.0	155.9 tr.t.
10.50	100.0	42.7	173.3 anh.
11.15	110.1	44.4	189.5
12.10	120.3	54.0	247.0

Lilich and Dzhurinskii, 1956

m	f.t.	m	f.t.
4.02	0	4.49	30
.09	5	.56	35
.16	10	.65	40
.24	15	.73	45
.32	20	.81	50
.41	25		

Lilich and Mogilev, 1956

m	pH	m	pH
	25°		
0	6.8-6.95	0.461	4.11
0.0155	6.19	0.946	3.54
.0444	5.74	1.47	3.08
.204	4.73	3.67	2.12

Hering and Leroy, 1939

mol%	t	d	mol%	t	d
7.17	8.7	1.739	15.8	126.1	2.115
7.47	20.0	.750	17.2	115.2	.190
7.73	30.0	.760	18.65	100.0	.251
8.02	40.0	.771	19.9	80.0	.321
8.33	50.0	.783	21.15	60.0	.380
8.67	60.0	.798	22.5	80.0	.407
9.05	70.0	.813	23.6	100.0	.419
9.46	80.0	.830	24.7	115.2	.435
9.94	90.0	.851	25.6	126.1	.447
10.50	100.0	.874	25.9	128.7	.451
11.15	110.1	.902	28.55	148.7	.496
12.10	120.3	.943	29.2	153.3	.519
			38.4	150.0	.676

Water + Cadmium sulfate ($CdSO_4$)

Heterogeneous equilibria .

Tammann, 1885

%	p
100°	
14.13	752.9
20.54	749.3
23.79	747.1
29.44	741.7
·33.30	736.8

Dieterici, 1923

m	p	m	p
	0°		
0	4.579	2.190	4.365
0.504	.555	2.500	.310
1.011	.513	3.179	.112
1.498	.463		

Ishikawa and Murooka, 1933

t	p
sat. sol. (8+3)	
20	15.65
25	21.17
30	28.44
35	37.84
40	49.68

Lescoeur, 1895

t	p dissoc. (1+1)	sat.sol.
14.6	-	9.2
20	-	12.9
60	-	155
84	-	290
165	55	-
170	115	-
174.5	162	-

Robinson and Stokes, 1949

m	osmotic coefficient	m .	osmotic coefficient
	25°		
0.1	0.565	1.0	0.452
0.2	0.513	1.2	0.461
0.3	0.490	1.4	0.476
0.4	0.476	1.6	0.496
0.5	0.466	1.8	0.522
0.6	0.458	2.0	0.551
0.7	0.452	2.5	0.632
0.8	0.450	3.0	0.726
0.9	0.449	3.5	0.832

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
4.36	100.11	24.77	100.62
9.89	100.12	26.88	100.72
13.11	100.29	29.22	100.84
17.12	100.36	32.14	101.01
19.85	100.39	34.84	101.19
21.73	100.51		

Dupuy, 1884

%	f.t.
41.82	15
46.80	30
51.51	60
54.26	100

Etard, 1894

%	f.t.	%	f.t.
35.9	0	41.6	94
37.5	10	27.7	130
41.5	24	14.7	165
42.0	30	7.1	188
49.7	65	2.3	200
43.5	86		

Kohnstamm and Cohen, 1898

%	f.t.	%	f.t.
	(8+3)		
43.02	0	43.21	15
43.07	5	43.23	16
43.09	7	43.22	17
43.13	9	43.22	18
43.16	11.5	43.24	19
43.19	13	43.44	25

Kahlenberg, 1901

%	f.t.
2.98	-0.313
7.92	-0.742
13.52	-1.322
18.46	-1.968
20.71	-2.330

Jones and Caldwell, 1901

c	f.t.
27.730	-2.526
20.797	-1.774
13.864	-1.175
11.089	-0.933
8.322	-0.714
5.513	-0.505
2.756	-0.263

Steinwehr, 1902

%	f.t.	%	f.t.
	(8+3)		
43.20	13.7	43.26	16.96
43.21	14.98	43.28	18.00
43.23	15.00	43.30	19.00
43.24	16.00	43.44	25.00

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.063	-0.201	0.625	-1.588
.125	-0.356	0.875	-2.388
.250	-0.658	1.000	-2.870
.500	-1.259	1.250	-4.160

Koppel, 1905

%	f.t.	
43.00	0	(8+3)
43.75	30	
44.01	40	

E= -16.8°

Cohen and Sinnige, 1909

P	% electric method	% direct method
	25°	
250	43.67	43.63
500	43.83	43.82
750	44.01	43.44
1000	44.11	44.06

Wolters, 1917 and Cohen and Wolters, 1920

%	f.t.	%	f.t.
	(8+3)		
43.047	- 3	43.008	- 9
43.020	- 6	43.028	-12

Benrath and Thönnessen, 1932

%	f.t.	
43.29	0	(8+3)
43.36	11.3	
43.73	25	
43.98	34	
44.09	36.5	
44.93	34.7	(1+1)I
44.48	40.5	
43.53	50	
42.10	64	
41.21	71	
40.95	74.5	tr.t.
40.97	75	(1+1)II
40.43	79	
39.98	86.3	
37.23	97	
36.85	99	

tr.t. (8+3) - (1+1)I : 44.32 % 41.5°

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
		(1+1)	
4.9	187	14.8	161
9.8	172	15.6	159
12.3	166	26.9	131
		32.3	119

Properties of phases .

Fouqué, 1867

%	d 0°	d 4°
1.38	1.0122	1.0121
5.42	1.0514	-
15.47	1.1552	1.1529

Grotrian, 1883

%	d	%	d
	18°		
1	1.0084	25	1.2950
5	.0486	30	.3725
10	.1026	35	.4575
15	.1607	36	.4743
20	.2245		

Jahn, 1891

c	d
	20°
0	0.9982
11.298	1.0888
22.114	1.1760

Schönrock, 1893

%	d
	16°
0	0.999
21.3671	1.24211
29.4654	1.36289

De Muynck, 1894

%	d^{18}	%	$d^{19.5}$
0	0.999	0	0.998
5.639	1.055	13.17	1.120
9.942	.101	29.07	.345
18.172	.200	46.83	.593
25.121	.297	54.93	.739
		74.50	2.338

Le Blanc and Rohland, 1896

%	d
20°	
0	0.99823
13.40	1.1409
16.79	1.1826

Barnes and Scott, 1898

%	d	%	d
18.2°			
0	0.9986	13.27	1.1437
0.464	1.0033	18.35	.2084
1.45	.0132	24.17	.2901
2.52	.0242	26.85	.3310
6.12	.0619	31.53	.4080
7.46	.0764	39.86	.5639
9.97	.1045		

Jones, 1904; Jones and Getman, 1904

M	d	M	d
0°			
0.063	1.015604	0.626	1.172224
.125	.032760	0.875	.239488
.250	.070824	1.000	.272040
.500	.139204	1.250	.337292

Irueste, 1915

%	d	$\alpha.10^6$	$\beta.10^7$
15°			
0	0.99913	-	-
1.43	1.01416	155	47
2.92	.02983	170	45
5.58	.06091	194	41
12.16	.13108	241	34
25.48	.30852	296	17

$d^t = d^{15}(1-\alpha(t-15)-\beta(t-15)^2)$

Cohen and Moesveld, 1920

%	d	%	d
19°			
0	0.99842	22.52	1.25926
1.89	1.01686	27.52	.33257
3.79	.03582	32.49	.41303
5.67	.05521	37.54	.50255
7.55	.07519	43.30	.61799
12.53	.13116	44.63	.64609
17.51	.19228		
30°			
0	0.99567	27.605	1.32892
7.9865	1.07658	28.323	.34000
8.619	.08344	32.029	.39991
15.336	.16079	37.092	.50031
16.389	.17373	42.063	.58587
23.140	.26499	43.564	.61760

Flöttmann, 1928

%	t	d
43.176	15	1.6159
43.42	20	.6165
43.62	25	.6186

Gibson, 1934

%	d	%	d
25°			
0	0.9971	20	1.221
10	1.100	30	.368
15	1.158	40	.525

Okazaki, 1935

%	d	%	d
28°			
7.80	1.0751	27.09	1.3202
15.17	.1589	33.23	.4187
20.49	.2265	40.18	.5456

Guillaume, 1946

%	d
20°	
11.65	1.1233
27.01	.3271
33.1	.426

Gibson, 1934

%	π (1-1000 bars)	%	π (1-1000 bars)
	25°		
0	36.46	20	24.99
10	31.95	30	18.85
15	28.39	40	13.91

Wagner, 1883

%	η (water°=1) 15°	25°	35°	45°
7.140	0.7890	0.6181	0.4990	0.4131
14.660	0.9617	0.7236	0.5808	0.4878
22.011	1.2081	0.9185	0.7346	0.6013

Jahn, 1891

c	n H_α	n D	n H_β
	20°		
0	1.3315	1.3332	1.3375
11.298	1.3435	1.3454	1.3497
22.114	1.3568	1.3613	1.3470

De Muynck, 1894

%	n_D 15°	n_D 20°	%	n_D 20°
0	1.33337	1.33298	0	1.33298
5.639	.34223	.36155	13.17	.35690
9.942	.34811	.34743	29.07	.39922
18.172	.36149	.36081	46.83	.44115
25.121	.37345	.37277	54.93	.46590
			74.50	.56805

Le Blanc and Rohland, 1896

%	n_D
20°	
0	1.3333
13.40	.3523
16.79	.3578

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.063	1.32790	0.625	1.34912
.125	.33037	0.875	.35799
.250	.33535	1.000	.36101
.500	.39466	1.250	.37028

Flöttmann, 1928

%	t	n_D
43.176	15	1.40999
43.42	20	.40952
43.62	25	.40920

Jahn, 1891

c	$(\alpha)_{magn.}$
20°	
0	1
11.298	0.47334
22.114	0.42163

Schönrock, 1893

%	$(\alpha)_{magn.}$
16°	
0	1
21.3671	0.3597
29.4654	0.3423

Guillaume, 1946

%	n 5780Å	$(\alpha)^*_{magn.}10^6$
	20°	
11.65	1.3516	3.679
27.01	.3783	.280
33.1	.3897	.117

* in radians, gauss, centim.

Okazaki, 1935

%	Verdet's constant.10^5 (3441 Å)
28°	
7.80	4508
15.17	4615
20.49	4706
27.09	4832
33.23	4920
40.18	5088

Grotrian, 1883

%	κ	τ
	18-22°	
1	41.0	0.0210
5	143	.0206
10	244	.0206
15	321	.0208
20	384	.0214
25	425	.0223
30	430	.0236
35	420	.0251
36	417	.0255

Jones and Caldwell, 1901

c	molecular conductivity	c	molecular conductivity
	25°		
27.761	32.5	0.277	115.2
13.880	45.5	0.139	130.0
6.940	57.1	0.069	165.2
2.776	71.8	0.027	190.5
1.388	84.4	0.014	209.2
0.694	99.8	0.007	225.7

Jones, 1904 and Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.063	67.10	0.625	34.20
0.125	56.30	0.875	31.51
0.250	46.97	1.000	27.07
0.500	38.46	1.250	20.40

Godlewski, 1902

c	λ	c	λ
		25°	
22.420	151.6	0.5605	320.6
11.210	185.0	0.2242	366.2
5.605	216.0	0.1121	398.4
2.242	254.4	0.05605	432.2
1.121	285.5	0.02242	474.0

Godlewski, 1902

%	e	%	e
0.01	0	7.41	0.0605
.104	0.0212	8.26	.0620
.25	.0260	9.45	.0638
.40	.0293	13.04	.0681
.69	.0343	17.45	.0722
1.03	.0380	20.00	.0744
.48	.0415	23.08	.0769
.96	.0444	24.35	.0779
2.44	.0467	27.54	.0804
3.38	.0504	30.30	.0828
3.96	.0524	31.51	.0842
4.76	.0548	33.33	.0866
5.66	.0570	35.55	.0901
6.54	.0589		

Cohen and Moeveld, 1920

%	U	%	U
	19°		
0	0.9992	22.52	7637
1.89	9763	27.52	7188
3.79	9544	32.49	6744
5.67	9339	37.54	6317
7.55	9120	43.30	5833
12.53	8612	44.64	5725

Water + Cadmium ammonium sulfate ($CdN_2H_8O_8S_2$)

Rüdorff, 1872

%	f.t.	%	f.t.
6.89	-1.15	17.49	-3.00
8.13	-1.35	20.06	-3.55
12.65	-2.15	21.65	-3.90
15.17	-2.50		

Jones and Caldwell, 1901

c	f.t.
20.203	-3.052
16.164	-2.493
12.111	-1.914
8.082	-1.345
4.041	-0.719
2.020	-0.378

Water + Cadmium sodium sulfate ($Na_2CdS_2O_8$)

Koppel, 1905

%	f.t.	%	f.t.
37.32	24	38.16	35
37.62	25	38.53	40
37.84	30	(4+1)	

Water + Cadmium potassium sulfate ($K_2CdS_2O_8$)

Wyrouboff, 1901

%	f.t.	%	f.t.
(2+1)		(3+2)	
30.06	16	29.83	26
31.89	31	29.97	31
32.15	40	30.28	40
		30.98	64

Water + Cadmium formate ($CdC_2H_2O_4$)

Ashton, Houston and Saylor, 1933

%	f.t.	%	f.t.
	(2+1)		
7.7	0	20.3	40
10.0	10	27.8	50
12.6	20	37.3	60
15.7	30	43.0	66
	anhydre		
43.5	70	46.0	90
44.7	80	48.6	100

Water + Cadmium acetate ($Cd\ C_4H_6O_4$)

Guillaume, 1946

%	d	$(\alpha)^*$ magn. 10^6 5780 Å	n
	20°		
13.30	1.1045	3.763	1.3518
21.62	.1602	.625	.3620
30.27	.2364	.489	.3743

* in radians, gauss, centim.

Water + Cadmium benzene sulfonate ($CdC_{12}H_{10}O_6S_2$)

Ephraim and Seger, 1925

c	f.t.	c	f.t.
27.969	18	55.764	64.5
35.511	34	69.200	80.5
44.936	49.5		

Water + Mercuric chloride ($HgCl_2$)

Kume, 1938

mol %	p	mol %	p
1.582	0.4625	31.715	7.803
3.464	0.9797	46.098	10.173
7.219	1.874	59.697	11.739
12.358	3.316	73.261	11.921
20.974	5.350	86.787	9.243

Beckmann, 1890

%	b.t.	%	b.t.
0	100	11.72	100.215
4.16	100.080	11.82	100.235
7.99	100.159	14.96	100.281
8.00	100.165	16.62	100.325
8.89	100.170	18.18	100.338

Baroni, 1893

%	b.t.	%	b.t.
3.56	100.075	13.36	100.267
6.53	100.125	17.21	100.327
9.83	100.189	20.22	100.395

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
3.23	100.06	25.87	100.50
7.90	100.15	31.45	100.61
13.40	100.25	34.47	100.65
19.92	100.38		

Hannay, 1873

%	f.t.	%	f.t.
4.09	0.0	16.23	60
4.35	2.2	18.42	65
4.61	4	20.81	70
4.78	5	22.70	75
4.85	10	26.52	80
5.59	15	28.74	85
5.93	20	32.30	90
6.68	25	36.71	95
7.17	30	37.82	96.5
8.86	35	41.56	97.5
9.72	40	52.51	99
10.61	45	55.73	100
12.37	50	56.02	100.8

Etard, 1894

%	f.t.	%	f.t.
3.9	+0.1	28.2	87
4.8	4.5	39.3	100
5.1	7.5	59.7	121
5.2	13.8	69.7	127
7.1	25.1	77.0	140
7.6	29.5	78.4	150
9.9	38.0	80.2	159
11.3	49	81.7	160
15.1	61	81.8	165
23.6	80		

Benrath, Gjedebo and al; 1937

%	f.t.	%	f.t.
38.9	105	82.6	164
49.0	116	86.5	175
55.3	123	88.8	182
61.5	129	91.4	195
65.1	133	93.0	206
70.6	141	94.8	223
73.2	145	96.0	235
80.0	157	100.0	275

Kume, 1938

mol%	f.t.	mol%	f.t.
1.582	80	31.715	180
3.464	100	46.098	200
7.219	120	59.697	220
12.358	140	73.261	240
20.974	160	86.787	260

Eddy and Menzies, 1940

m	f.t.	m	f.t.
0.173	0.9	0.995	75.4
.244	20.93	1.612	91.6
.380	39.25	2.110	99.7
.566	55.85	2.773	105.6
.687	62.7		

Mendelejeff, 1860

%	d	%	d
	20°		
0	0.998	10	1.095
8	1.071	11	.1035
9	1.0815	12	.115
		13	.127

Hannay, 1873

%	t	d	%	t	d
4.09	0.0	1.039	16.23	60	1.116
4.35	2.2	.039	18.42	65	.124
4.61	4	.039	20.81	70	.146
4.78	5	.041	22.70	75	.165
4.85	10	.042	26.52	80	.181
5.59	15	.045	28.74	85	.203
5.93	20	.049	32.30	90	.225
6.68	25	.056	36.71	95	.272
7.17	30	.060	37.82	96.5	.287
8.86	35	.067	41.56	97.5	.310
9.72	40	.076	52.51	99	.346
10.61	45	.083	55.73	100	.354
12.37	50	.089	56.02	100.8	.357

Mc Clung and Mc Intosh, 1902

% X-ray absorption at room temp.	d
51.1	1.000
58.1	.002
68.8	.004
79.8	.011
89.9	.026
96.6	.052

Water + Mercuric bromide ($HgBr_2$)

Benrath, Gjedebo and al; 1937

%	f.t.	%	f.t.
12.0	142	80.2	188
22.4	164	88.1	189
33.5	173	92.8	193
59.1	185	96.2	201
74.7	187	100	237

Water + Mercuric iodide (HgI_2)

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
3.7	196	48.0	324
9.5	229	49.7	326
9.8–11.5	241	53.8	328
10.0	255	54.4	328
12.0°	243	65.5	334
15.5	256	68.0	336
15.7	257	74.3	337
18.7	272	75.0	338
23.7	285	82.7	336
27.5	295	90.5	322
30.5	300	95.3	293
41.0	314	97.4	262

Water + Mercuric cyanide (HgC_2N_2)

Tammann, 1885

%	p
100°	
8.43	755.2
14.06	751.5
16.23	750.7
20.93	747.2
23.41	745.0

Corbet, 1926

10 % f.t. = 25°

Benrath, Gjedebo and al; 1937

%	f.t.	%	f.t.
35.8	108	56.7	156
42.9	125	63.4	175
49.8	140	72.8	209

Herz and Martin, 1923

t	d	η	t	d	η
30.321%					
20	1.2206	2040	60	1.1779	1030
30	.2099	1640	70	.1670	900
40	.1993	1390	80	.1560	799
50	.1886	1200	90	.1453	717

Water + Mercuric perchlorate ($HgCl_2O_8$)

Lilich and Dzhurinskii, 1956

m	f.t.	m	f.t.
6.07	0	6.87	20
6.30	5	7.45	25
6.49	10	8.09	30
6.70	15	8.70	35

Geffcken, 1929

m	d	n_{He}	m	d	n_{He}
25°					
0	0.99707	1.33259	5.8625	1.76315	1.40237
1.0762	1.16533	.34826	9.4425	2.09353	.43176
3.1071	.44606	.37332	11.202	2.22960	.44389

Lilich and Mogilev, 1956

m	pH	m	pH
25°			
0	6.8-6.95	0.0458	2.40
0.0000817	4.43	.0840	2.22
.000311	3.74	.157	2.04
.00117	3.28	.576	1.64
.00172	3.14	1.08	1.13
.0045	2.94	2.32	0.30
.0187	2.63	5.51	-1.56

Water + Mercuric acetate ($HgC_4H_6O_4$)

Kanonnikoff, 1885

%	d	n H_α	n D	n H_β
22°				
0	0.9978	1.33110	1.33278	1.33718
21.03	1.1656	.351594	.35400	.35870

LIII. WATER + SALTS OF Cu, Mn, Fe, Ni and Co .

Water + Cupric chloride ($CuCl_2$)

Ewan and Ormandy, 1892

%	t	p	%	t	p
1.66	21.023	18.410	2.03	18.000	15.219
3.13	21.010	18.282	3.60	17.950	15.058
5.30	21.005	18.092	6.86	17.930	14.842
7.45	20.967	17.741	9.66	17.950	14.621
11.49	20.920	17.472	12.79	17.983	14.404
16.80	20.930	16.868	15.24	18.198	14.335
19.87	20.900	16.347	24.63	18.190	13.291
22.72	21.040	16.162	30.73	18.190	12.470
30.54	20.970	14.946			
2.03	15.770	13.206	17.92	14.425	11.078
4.79	15.797	13.081	23.40	14.222	10.425
9.90	15.920	12.859	28.48	14.267	9.938
14.27	15.983	12.574	40.12	14.808	9.028
19.92	16.060	12.115	42.83	14.354	8.197
24.23	16.032	11.647			
29.07	16.097	11.168			
33.81	16.310	10.538			

Lescoeur, 1894

t	p	t	p
	sat.sol.		
10	4.8	65	159.0
20	9.8	70	198.0
40	41.5	80	301.0
60	125.0		

t	p
(8+1)	
80	62
90	107
100	185

Speranski, 1911

t	p	t	p
	sat.sol.		
22.4	13.70	35.95	29.71
25.0	15.97	37.8	32.75
28.35	19.42	38.4	33.86
30.35	21.76	40.45	37.78

Derby and Yngve, 1916

t	p sat.sol. (2+1)	diss. (2+1)-(1+1)
17.90	-	3.7
26.60	23.3	8.7
31.40	30.9	11.7
39.88	49.5	18.8
53.54	94.3	54.3
64.18	-	105.6
64.45	150.6	-

Boswell and Cantelo, 1922

N	Dp/po	N	Dp/po
	23°		
7.000	0.214	2.000	0.048
5.000	.140	1.500	.035
3.500	.094	1.000	.018

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.845	1.4	1.000
0.2	0.843	1.6	1.022
0.3	0.848	1.8	1.043
0.4	0.860	2.0	1.062
0.5	0.876	2.5	1.100
0.6	0.892	3.0	1.131
0.7	0.908	3.5	1.160
0.8	0.922	4.0	1.183
0.9	0.938	4.5	1.201
1.0	0.952	5.0	1.219
1.2	0.978	5.8	1.249

Robinson and Stokes, 1940

Isopiestic solutions

m_1	m_2	m_1	m_2
0.1085	0.1495	0.1128	0.1558
.2516	0.3521	.3998	0.5750
.7236	1.107	.8388	1.300
1.206	1.945	1.259	2.035
1.481	2.430	1.553	2.574
2.111	3.577	2.415	4.121
2.769	4.729		
0.2048	0.2838	0.2474	0.3462
.5334	0.7856	0.6555	0.9860
.9378	1.472	1.054	1.677
1.297	2.105	.328	2.157
1.847	3.106	.996	3.366
2.530	4.314	2.652	4.522

m_1 = cupric chloride, m_2 = potassium chloride

Rüdorff, 1862

%	f.t.
2.33	-0.85
4.68	-1.8
12.99	-6.6
14.18	-7.5

Reicher and van Deventer, 1890

%	f.t.
43.06	17
44.7	31.5

Ewan and Ormandy, 1892

%	f.t.
4.76	-1.905
9.09	-4.12
16.67	-9.63

Etard, 1894

%	f.t.	%	f.t.
37.0	-20	44.0	39
38.8	-5	55	46.5
39.3	+12	68	47.9
41.7	17	73	48.6
43.2	32	91	51.0

Jones and Getman, 1904; Jones 1904; Jones and Bassett, 1905

M	f.t.	M	f.t.
0.0651	-0.331	1.8210	-12.960
.1301	-0.639	2.0816	-15.480
.2602	-1.273	2.6020	-20.820
.5204	-2.711	3.0000	-25.500
.7806	-4.413	3.5000	-31.500
1.3010	-8.395	4.3710	-44.500
1.5612	-10.383		

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.01	-0.05703	0.50	-2.669
.05	.24944	0.75	-4.245
.075	.37075	1.00	-5.994
.100	.48665	1.50	-10.105
.25	-1.2237	2.00	-15.294

Schreinemakers, 1910

mol %	f.t.
(2+1)	
8.934	15
9.305	25
9.689	35

Boye, 1933

%	f.t.		%	f.t.	
0	0		42.6	23.0	
7.8	-2.9		43.3	24.9	
11.3	-4.5		43.5	25.5	
16.6	-9.2		43.6	25.7	tr.t.
17.5	-10.0		43.6	26.0	(2+1)
24.5	-17.5		43.6	26.8	
25.4	-18.1		43.6	28.5	
34.1	-31.4		43.6	31.0	
35.4	-34.0		43.7	33.9	
37.4	-37.3		43.9	35.8	
37.5	-37.5		44.3	38.4	
37.6	-38		44.7	39.9	
39.9	-43.4	E	44.8	41.0	
39.5	-31.4	(4+1)	45.0	41.6	
39.4	-29.9		45.2	42.2	tr.t.
39.4	-26.8		45.1	42.7	(1+1)
40.7	0		45.0	43.1	
40.9	0.8		45.0	45.3	
41.5	10.0		44.9	46.1	
41.6	11.5		44.9	46.8	
41.8	13.0		45.0	50.0	
41.9	14.0		45.7	55.3	
42.1	15.0	tr.t.	46.7	61.2	
42.1	15.5	(3+1)	47.8	68.3	
42.1	16.0		48.5	73.0	
42.1	17.0		48.7	75.1	
42.1	18.0		49.8	82.0	
42.1	18.5		50.9	89.9	
42.1	19.0		52.0	96.1	
42.2	20.0		52.8	101.8	
42.3	21.6		55.0	116.8	

Kamecki and Trau, 1955

tr.t. (2+1) - anh. 131.8°

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987		
1	1.0078	21	1.2346
2	.0169	22	.2484
3	.0260	23	.2623
4	.0351	24	.2762
5	.0442	25	.2901
6	.0534	26	.3041
7	.0627	27	.3181
8	.0720	28	.3321
9	.0813	29	.3461
10	.0906	30	.3600
11	.1035	31	.3766
12	.1164	32	.3932
13	.1293	33	.4098
14	.1422	34	.4269
15	.1550	35	.4429
16	.1680	36	.4596
17	.1811	37	.4763
18	.1942	38	.4930
19	.2073	39	.5097
20	.2207	40	.5265

Favre and Valson, 1874

m	d	m	d
	22.9°		
0	0.998	3.0	1.299
0.5	1.059	3.5	.340
1.0	.114	4.0	.379
1.5	.165	4.5	.416
2.0	.213	5.0	.453
2.5	.257		

Charpy, 1893

%	d				
	0°	7°	30°	49.2°	65.4°
0.0000	0.9999	0.9999	0.9955	0.9884	0.9806
8.0732	1.0795	1.0784	1.0734	1.0612	1.0519
14.5820	.1493	.1468	.1385	.1271	.1164
20.6697	.2202	.2169	.2064	.1928	.1820
26.1129	.2879	.2846	.2721	.2569	.2452
30.9255	.3527	.3486	.3351	.3191	.3054
35.3839	.4171	.4125	.3972	.3800	.3662
39.4170	.4795	.4744	.4579	.4396	.4251

Holland, 1898

%	d
18°	
1.40	1.0111
10.30	.0996
19.78	.2057
29.24	.3276
36.51	.4328

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	d	M	d
	0°		
0.0650	1.007940	1.8210	1.209400
.1301	.015736	2.0816	.239996
.2602	.031764	2.6020	.296360
.5204	.063652	3.0000	.335796
.7806	.093832	3.5000	.391508
1.3010	.153120	4.3710	.479236
.5612	.183364		

Chéneveau, 1907

%	d	%	d
	16°		
0	0.9990	14.13	1.1408
5.12	1.0484	18.11	.1858
9.22	.0956	19.99	.2072
12.02	.1184	21.80	.2290

Jones and Pearce, 1907

M	d	M	d
	20°		
0	0.998230	0.50	1.049615
0.01	0.999433	0.75	.088978
.05	1.004586	1.00	.118264
.075	.007475	1.50	.175531
.100	.010829	2.00	.232364
.250	.029163		

Schneider, 1910

N	d	N	d
	18°		
0	0.99862	1.077	1.06146
.1077	1.00535	2.154	.1215
.2154	.01139	4.308	.2417
.538	.03142		

Heydweiller, 1909 and 1912

N	d
18°	
0.1	1.00487
0.2	1.01098
0.5	1.02948
1.0	1.05943
2.0	1.1175
4.0	1.2291

Clausen, 1912

N	d 6°	d 18°	d 30°
0.490	1.03075	1.02887	1.02560
0.982	.06070	.05826	.05456
1.965	.11894	.11690	.11561
3.975	.23288	.22912	.22650

Herz, 1914

N	d
25°	
0	0.9971
0.733	1.0400
1.466	1.0827
2.932	1.1650
4.398	1.2440
5.864	1.3205

Irueste, 1915

%	d	$\alpha .10^6$	$\beta .10^7$
15°			
0	0.99913	-	-
3.59	1.03366	190	45
7.53	.06675	254	34
11.32	.11000	285	28
23.69	.25422	402	15
32.52	.36578	463	11

$$d^t = d^{15}(1-\alpha(t-15)-\beta(t-15)^2)$$

Herz, 1917

N	d	N	d
25°			
0	0.9971	2.932	1.1649
0.733	1.0404	4.398	.2444
1.466	1.0831	5.864	.3208

Berthier, Courty and Gauthier, 1952

%	t	d	%	t	d
0.122	22	1.0010	15.60	25	1.1512
.213	22.5	.0013	16.99	25	.1673
.409	23	.0030	18.65	25.5	.1862
.612	23	.0056	20.45	25	.2061
.836	23.5	.0052	21.75	25	.2211
.938	19	.0063	21.97	28	.2269
1.34	19	.0104	22.64	23	.2311
1.93	19.5	.0170	23.68	28	.2438
2.77	19	.0240	24.83	27	.2571
3.66	19	.0328	26.70	25.5	.2828
4.81	19	.0439	28.78	26	.3131
5.85	25	.0519	31.10	25	.3423
6.50	25	.0591	31.45	23	.3484
6.93	25.5	.0649	31.61	24	.3467
7.27	26	.0669	31.95	24	.3577
7.75	25.5	.0729	33.65	23	.3832
8.71	25.5	.0814	35.99	15.5	.4156
9.57	25	.0886	36.73	21	.4419
10.44	25	.0992	36.77	15.5	.4562
11.16	24	.1108	38.94	15.5	.4617
12.09	25	.1148	39.29	15	.4624
13.09	25	.1172	40.88	15	.5958
13.50	26	.1300	42.37	16	.5181
14.91	26	.1447			

Wagner, 1883

%	η (water°=1) 10°	15°	20°	25°	35°	45°
21.349	1.3800	1.1740	1.0080	0.9585	0.7703	0.6325
33.027	2.0910	1.7840	1.5240	1.3740	1.0760	0.8708

Schneider, 1910

N	η	N	η
18°			
0.1077	1058	1.077	1290
.2154	1083	2.154	1553
.5380	1109	4.308	2226

Herz, 1914 and 1917

N	η	N	η
25°			
0	895	2.932	1508
0.733	1047	4.398	1900
1.466	1180	5.864	2404

Mazetti, 1924

N(18°)	η 13.2-13.7°	34.0-34.6°	45.4-45.8°	59°
0.251	1278	821	662	516
.501	1306	853	692	540
.752	1346	881	725	570
1.254	1476	951	766	610
1.504	1515	975	796	636
1.755	1574	1004	819	654
2.006	1648	1014	850	675
2.507	1835	1138	941	748
7.631	5983	3303	2548	1911

de Heen, 1883

Surface tension (see author)

Walter, 1889

%	n_D	%	n_D
15°			
0	1.3334	12.76	1.3671
1.81	.3401	15.87	.3766
2.52	.3417	19.0	.3865
5.17	.3479	26.7	.4115
10.52	.3601	31.6	.4283
11.09	.3619	38.2	.4549

Jones and Getman, 1904
and Jones and Bassett, 1905

M	n_D	M	n_D
0°			
0.0651	1.32735	1.3010	1.36163
0.1301	.32989	1.5612	.36934
0.2602	.33428	1.8210	.37579
0.5204	.34106	2.0816	.38222
0.7806	.34842	2.6020	.39491

Chéneveau, 1907

%	n_D	%	n_D
16°			
0	1.3334	14.13	1.3665
5.12	.3446	18.11	.3769
9.22	.3558	19.99	.3820
12.02	.3612	21.80	.3868

Holland, 1898

t	κ	$\tau.10^4$
36.51 %		
10	561.7	234
18	691.3	234
20	724.3	248
30	896.9	248
40	1075.4	252
50	1256.9	255
60	1427.5	253
70	1596.8	252
80	1756.8	248
90	1905.5	244
29.24 %		
10	725.8	205
18	868.2	205
20	905.7	214
30	1091.2	214
40	1273.7	212
50	1455.1	211
60	1630.5	209
70	1791.0	204
80	1791.0	199
90	2066.2	192
19.78 %		
10	814.7	186
18	957.0	186
20	994.2	192
30	1177.2	192
40	1353.1	188
50	1525.8	186
60	1682.4	180
70	1824.2	174
80	1948.0	167
90	2055.5	159
10.3 %		
10	628.7	190
18	741.6	190
20	770.5	192
30	912.0	192
40	1052.1	190
50	1186.0	187
60	1305.4	181
70	1414.3	174
80	1515.6	168
90	1597.7	160
1.40 %		
10	124.8	209
18	149.9	209
20	156.6	221
30	189.7	221
40	223.6	224
50	259.5	228
60	293.6	228
70	325.2	225
80	356.4	222
90	383.7	217

Jones and Getman, 1904

Jones, 1904 and Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.0650	94.42	1.8210	35.02
0.1301	88.06	2.0816	30.81
0.2602	78.38	2.6020	22.58
0.5204	67.68	3.0000	17.48
0.7806	59.30	3.5000	13.05
1.3010	48.22	4.3710	7.61
1.5612	39.32		

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.010	110.55	0.50	69.08
0.050	95.88	0.75	61.03
0.075	93.27	1.00	54.01
0.10	90.21	1.50	41.81
0.25	80.20	2.00	31.56

Heydweiller, 1909 and 1912

N	$\varkappa$	N	$\varkappa$
	18°		
0.1	83.3	1.0	569
0.2	154.4	2.0	870
0.5	335.5	4.0	1048

Clausen, 1912

N	λ 6°	18°	30°
0.490	50.73	67.22	84.18
0.982	43.68	57.16	71.02
1.965	33.83	43.83	53.84
3.975	20.08	26.03	32.02

Mazetti, 1924

N (18°)	$\varkappa$ 13.4-14.0°	34.4-34.6°	45.6-45.9°	59.0-59.1°
0.251	176.8	265.4	314.2	367.6
0.501	321.8	477.5	567.3	662.8
0.752	455.3	667.0	783.1	909.5
1.254	676.9	982.3	1152.0	1345.0
1.504	769.6	1116.2	1308.0	1524.0
1.755	852.1	1230.5	1439.5	1679.2
2.006	923.0	1333.7	1559.8	1813.2
2.507	1038.2	1496.5	1747.6	2037.4
7.631	856.1	1326.5	1589.9	1902.5

Chéneveau, 1910

%	χ
23°	
0	-0.79
2.80	-0.422
9.49	+0.500
15.73	+1.359

Cabrera and Moles, 1915

%	χ_{mol}	%	χ_{mol}
	20°		
3.02	15770	13.28	15740
5.03	15830	26.10	15630
8.57	15870	36.65	15660

Jacobson, 1916

%	χ
0.443	-0.665
1.907	-0.484
12.41	+0.805
24.41	+2.334
35.40	+3.588

Berthier, Courty and Gauthier, 1952

%	t	χ	%	t	χ
0.122	22	-0.676	15.60	25	1.137
.213	22.5	.673	16.99	25	.287
.409	23	.646	18.65	25.5	.474
.612	23	.608	20.46	25	.692
.836	23.5	.579	21.75	28	.811
.938	19	.595	21.97	23	.893
1.34	19	.538	22.64	28	.892
1.93	19.5	.471	23.68	28	2.033
2.77	19	.379	24.83	27	.174
3.66	19	.278	26.70	25.5	.416
4.81	19	.087	28.78	26	.661
5.85	25	.017	31.10	25	.944
6.50	25	+0.073	31.45	23	.992
6.93	25.5	.134	31.61	24	.966
7.27	26	.165	31.95	24	3.038
7.75	25.5	.220	33.65	23	.379
8.71	25.5	.333	35.99	15.5	.619
9.57	25	.425	36.73	21	.714
10.44	25	.524	36.77	15.5	.897
11.16	24	.583	38.94	15.5	.946
12.09	25	.735	39.29	15	.912
13.09	25	.895	40.88	15	4.1696
13.50	26	.901	42.37	16	4.2813
14.91	26	1.040			

Mc Clung and Mc Intosh, 1902

d	% of X-ray absorption	d	% of X-ray absorption
	at room t.		
1.169	92.3	1.020	69.0
1.085	86.2	1.009	64.3
1.042	77.00	1.000	58.9
1.009	78.1	1.002	65.2
1.004	70.3	1.000	57.7

Marignac, 1876

%	U	%	U
	19.51°		
3.59	0.9563	22.95	0.7790
6.93	0.9200	42.65	0.6241
12.97	0.8642		

Biltz, 1899

t	heat conductivity	d
8-14°	425	1.255
28-36°	566	-
8-14°	407	1.125
28-36°	6665	-
8-14°	413	water
28-36°	662	-

Water + Cupric bromide ($CuBr_2$)

Carter and Megson, 1928

%	f.t.	%	f.t.
51.8	0.0	55.8	25.0
52.8	5.75	56.1	30.1
53.7	9.9	56.0	34.8
55.0	15.0	56.8	50.0
55.9	20.0	55.7	0.0

tr.t. (4+1) - anh. 17.97° metastable

Water + Potassium copper chloride (K_2CuCl_4)

Jones, 1904; Jones and Bassett, 1905

N	f.t.	d	λ
		0°	
0.05	-0.623	1.010500	228.1
.1	-1.200	.020900	212.0
.2	-2.425	.042400	196.5
.3	-3.668	.063300	186.0
.4	-5.00	.084300	183.6
.6	-7.50	.124700	159.0
.8	-10.00	.164400	143.3
1.0	-13.00	.203500	127.0

Water + Cupric chlorate ($CuCl_2O_6$)

Meusser, 1902

%	f.t.	%	f.t.
30.53	-12	62.17	18
39.14	-25	66.17	45
54.59	-31 (4+1)	69.42	59.6
57.12	-21	76.90	71
58.51	0.8		

d sat.sol. 18° = 1.695

Traube, 1895

%	d	%	d
		15°	
0	0.99913	6.945	1.05714
2.106	1.01620	10.016	.08444
4.778	1.03857	14.387	.12531

Rubien, 1911

N	d	N	d
		18°	
0	0.99862	0.9776	1.08776
.09948	1.00790	1.954	.17461
.1974	.01694	3.913	.34447
.4898	.04370		

Heydweiller, 1912

N	d	N	d
		18°	
0.0901	1.0070	0.901	1.0810
.1803	.0154	1.803	.1615
.451	.0400	3.605	.3186

Heydweiller, 1909

N	n_D	N	n_D
		18°	
0	1.33327	1.0	1.34835
0.1	.33482	2.0	.36278
0.2	.33634	4.0	.39013
0.5	.34087		

Rubien, 1911

N	n_D
18°	
0	1.33327
0.09948	.33481
0.1974	.33630
0.4898	.34072
0.9776	.34801
1.954	.36211
3.913	.38893

Heydweiller, 1912

N	κ
18°	
0.0901	69.1
0.1803	128.8
0.451	287.8
0.901	506.3
1.803	822.6
3.605	1077

Jacobson, 1916

%	χ	t
0.565	-0.679	82.5
1.334	-0.662	3.8
8.58	-0.099	88.95
51.74	+3.009	3.55

Water + Cupric nitrate (CuN_2O_6)

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
25°			
0.1	0.847	1.6	1.131
0.2	0.849	1.8	1.177
0.3	0.860	2.0	1.224
0.4	0.875	2.5	1.339
0.5	0.895	3.0	1.480
0.6	0.914	3.5	1.610
0.7	0.934	4.0	1.732
0.8	0.955	4.5	1.841
0.9	0.978	5.0	1.940
1.0	1.001	6.0	2.125
1.2	1.046	7.0	2.286
1.4	1.087	8.0	2.424

Rüdorff, 1872

%	f.t.	%	f.t.
3.58	-0.90	16.27	-6.27
7.96	-2.35	21.89	-9.30
12.97	-4.25		

Etard, 1894

%	f.t.	%	f.t.
38.8	-10	48.5	+ 8
41.6	- 3	54.1	+20
44.5	+ 3	61.2	+32

Funk, 1899

%	f.t.	%	f.t.
30.93	-16	63.39	26.4
34.29	-20	60.01	25 (3+1)
36.08	-23 (9+1)	60.44	30
37.38	-21	61.51	40
40.92	-20	62.62	50
39.52	-21 (6+1)	64.17	60
42.08	-10	65.79	70
45.00	0	67.51	80
48.79	+10	77.59	114.5
55.58	+20		

Jones and Getman, 1904
and Jones, 1904 and Jones and Bassett, 1905

M	f.t.	M	f.t.
0.0591	-0.326	1.1815	- 7.799
0.1182	-0.595	1.6541	-12.650
0.2362	-1.210	1.8904	-15.490
0.4726	-2.544	2.3630	-21.890
0.9452	-5.903		

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.01	-0.05736	0.50	-2.589
.025	-0.13852	.75	-4.190
.050	-0.25540	.935	-5.512
.075	-0.36979	1.50	-10.284
.25	-1.221	2.00	-16.89

Kasantsev, 1923

%	f.t.	%	f.t.
		(6+1)	
46.3	0	57.6	22.5
50.3	10	58.8	24.0
51.5	12.5	59.7	24.5
52.4	15	60.1	25
54.1	17.5	60.6	25.5
55.4	20		
		(3+1)	
60.2	22.5	63.5	50
60.3	25	67.3	75
60.9	30	71.5	100

tr.t. = 25.4°

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987	23	1.2384
1	1.0077	24	.2504
2	.0167	25	.2627
3	.0257	26	.2758
4	.0347	27	.2889
5	.0439	28	.3040
6	.0536	29	.3151
7	.0634	30	.3287
8	.0732	31	.3415
9	.0830	32	.3557
10	.0928	33	.3690
11	.1028	34	.3823
12	.1128	35	.3956
13	.1228	36	.4106
14	.1328	37	.4256
15	.1428	38	.4406
16	.1545	39	.4555
17	.1664	40	.4705
18	.1783	41	.4875
19	.1902	42	.5045
20	.2021	43	.5215
21	.2142	44	.5385
22	.2263	45	.5556

Long, 1880

%	d	%	d
	15°		
5	1.043	20	1.193
10	.089	25	.248
15	.139	35	.377

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	d	M	d
	0°		
0.0591	1.005824	1.1815	1.173156
.1182	.015404	.6541	.240192
.2362	.033456	.8904	.276736
.4726	.069060	2.3630	.345560
.9452	.141136		

Chéneveau, 1907

%	d	%	d
	16°		
0	0.9990	13.33	1.1198
4.78	1.0393	17.17	.1595
9.21	.0804	18.99	.1791
11.31	.1002	20.75	.1977

Jones and Pearce, 1907

M	d	M	d
	20°		
0	0.99823	0.5	1.07532
0.01	0.999728	.75	.11271
.025	1.002296	.935	.14060
.05	.006072	1.50	.22400
.075	.009983	2.00	.29032
.25	.03844		

Rubien, 1911

N	d	N	d
	18°		
0	0.99862	1.025	1.07757
.0995	1.00649	1.987	.14973
.1992	.01430	3.97	.29369
.4995	.03738		

Irueste, 1915

%	d	α .10^6	β .10^7
	15°		
0	0.99913	-	-
1.41	1.01146	172	56
4.23	.03427	209	42
12.33	.11206	317	27
19.65	.18333	388	17
31.68	.32369	491	5

$d^t = d^{15}(1-\alpha(t-15)-\beta(t-15)^2)$

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
12.944	1.1028
25.888	1.2049

Pesce, 1935

N	d
25°	
0.907	1.06559
1.886	.13749
2.891	.20996
4.760	.34237
7.003	.49449

Amiel, 1941

M	d	M	d
	20°		
0.1	1.0137	1	1.1514
.2	.0294	1.2	.1798
.3	.0453	1.5	.2240
.4	.0604	2	.2974
.5	.0753	2.5	.3659
.6	.0912	3	.4359
.8	.1207	4	.5702

Berthier, Courty and Gauthier, 1952

%	t	d	%	t	d
0.074	23	0.99806	26.989	19.5	1.26749
.155	23.5	0.99822	28.722	20	.28679
.334	19.5	1.00085	30.155	20	.30492
.778	19	.00434	31.474	19.5	.32313
1.207	18	.00752	33.754	19	.35258
1.809	20	.01209	34.546	21	.36366
2.929	19	.02294	35.929	20	.38056
4.081	22	.03162	36.079	21	.38311
5.248	22	.04215	37.480	21	.40086
6.609	22	.05409	38.850	21	.42403
7.288	21	.05967	40.409	21	.44046
7.938	21.5	.06610	41.697	20.5	.46323
8.394	21.5	.07780	42.879	20.5	.47989
10.119	22	.08601	43.661	20	.49198
12.963	21	.11399	45.013	20	.41833
15.482	21.5	.13801	45.498	21	.51709
17.150	20	.15591	46.057	20	.53150
17.146	21	.15668	47.527	21.5	.55165
18.892	21	.17342	48.363	20.5	.56769
20.700	22	.19535	50.706	23	.60326
23.063	20	.22116	52.582	23.5	.63301
25.532	21	.24595	54.401	24	.67052

Wagner, 1883

%	η(water=1) 15°	25°	35°	45°
18.99	0.9728	0.7604	0.6153	0.5129
26.68	1.2621	0.9883	0.8091	0.6856
36.71	3.8288	2.8383	2.1531	1.7224

Walter, 1889

%	n_D	%	n_D
	15°		
0	1.3334	18.4	1.3707
2.32	.3400	29.0	.3947
5.32	.3459	35.2	.4176
10.0	.3539	45.3	.4369
14.9	.3631		

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	n_D	M	n_D
	0°		
0.0591	1.32751	0.4726	1.33979
.1182	.32949	0.9452	.35329
.2362	.33199	1.1815	.35961

Cheneveau, 1907

%	n_D	%	n_D
	16°		
0	1.3334	13.33	1.3550
4.78	.3406	17.17	.3620
9.21	.3480	18.99	.3656
11.31	.3515	20.75	.3690

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.34756
.1	.33473	2.0	.36138
.2	.33619	4.0	.38756
.5	.34050		

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	1.025	1.34793
.0995	.33472	1.987	.36120
.1992	.33618	3.97	.38706
.4995	.34049		

Long, 1880

%	κ	τ	%	κ	τ
		18°			
5	346	0.0221	20	967	0.0205
10	605	.0215	25	1035	.0216
15	816	.0206	35	1009	.0237

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.0591	99.36	1.1815	52.34
.1182	93.05	.6541	40.70
.2362	82.97	.8904	35.84
.4726	73.16	2.3630	26.70
.9452	58.45		

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.010	108.84	0.75	62.0
0.025	101.2	0.93	56.89
0.050	95.7	1.5	44.0
0.075	91.86	2.0	33.47
0.25	79.4		

Heydweiller, 1912

N	κ	N	κ
	18°		
0.0996	82.7	1.987	910.3
.1992	151.7	3.971	1131
.4995	336.7	9.872	462.7
1.025	590.1		

Cheneveau, 1910

%	χ
24°	
7.54	-0.043
19.35	+1.108
25.10	+1.720

Sloan, 1910

M	λ	M	λ
	0°		
7.69	5.65	1.921	30.33
6.94	7.42	0.961	37.90
5.38	12.26	.480	43.20
3.843	18.96	.0883	47.00

Cabrera and Moles, 1915

%	χ^{mol}_{Cu}	%	χ^{mol}_{Cu}
	20°		
4.84	1597	29.75	1568
7.24	1589	35.71	1570
11.90	1584	43.79	1571
16.83	1577	50.90	1450
23.80	1577		

Amiel, 1941

M	χ	M	χ
	20°		
0.1	-0.546	1	+0.792
.2	-0.382	1.2	1.046
.3	-0.221	1.5	.413
.4	-0.065	2	.994
.5	+0.089	2.5	2.521
.6	+0.240	3	2.959
.8	+0.516	4	3.733
		satd (3+1)	+6.5

Berthier, Courty and Gauthier, 1952

%	t	χ	%	t	χ
0.074	23	-0.682	26.989	19.5	+1.664
.155	23.5	.677	28.722	20	.805
.334	19.5	.663	30.155	20	.944
.778	19	.627	31.474	19.5	2.107
1.207	18	.591	33.754	19	.247
1.809	20	.540	34.546	21	.321
2.929	19	.439	35.929	20	.401
4.081	22	.248	36.079	21	.440
5.248	22	.206	37.480	21	.557
6.609	22	.120	38.850	21	.656
7.288	21	.052	40.409	21	.776
7.938	21.5	+0.012	41.697	20.5	.930
8.384	21.5	.111	42.879	20.5	3.010
10.119	22	.186	43.661	20	.112
12.963	21	.417	45.013	20	.217
15.482	21.5	.671	45.498	21	.237
17.150	20	.811	46.057	20	.299
17.146	21	.901	47.527	21.5	.410
18.892	21	1.004	48.363	20.5	.496
20.700	22	.154	50.706	23	.673
23.063	20	.305	52.582	23.5	.837
25.532	21	.520	54.401	24	.968

Da Silveira, 1939

Raman spectra

Biltz, 1899

$d^{20°}$	heat conductivity	t
1.197	423	8-14°
"	662	28-36°
1.455	404	8-14°
"	563	28-36°
water	413	8-14°
"	662	28-36°

Water + Cupric perchlorate ($CuCl_2O_8$)

Smeets, 1931

%	f.t.	%	f.t.
0	0	51.5	-20.0
16.7	-5.0	54.2	0.0
25.4	-10.4	57.4	+25.0
33.8	-18.3	62.0	54.5
40.0	-29	65.0	69.0
45.0	-41	70.8	82.3
47.0	-50.5	71.2	86.0
48.0	-46	71.5	100.0
50.0	-33	71.8	114

E: -50.5°

Berthier, Courty and Gauthier, 1954

%	t	d	χ
0.236	20	1.000944	-0.6599
.373	24	.001197	.5901
.756	16	.005386	.5940
.790	20	.003808	.6195
2.161	20	.013822	.5259
2.2125	18	.015985	.5615
3.992	20	.031743	.4325
5.351	19	.035321	.3860
6.094	23	.046870	.3468
7.805	19	.059656	.1944
10.632	25	.082768	.0747
12.641	25	.102449	+0.0875
16.022	19	.135451	.3199
21.102	19.5	.185594	.6927
24.455	21	.220273	.9036
27.160	20	.254005	1.0808
28.150	24	.260508	.1294
30.417	26	.281700	.2511
30.597	23	.288244	.2828
33.460	23	.319993	.4380
34.135	24	.329040	.5156
36.177	24	.351181	.6499
38.881	25	.392091	.8586
39.851	21	.392950	.8415
41.690	25.5	.431484	2.0454
42.927	22	.448260	.0887
43.285	22	.454445	.0347
45.016	25	.477739	.2093
47.006	21	.511243	.3331
47.431	22	.520434	.3932
48.217	21	.529448	.4237
49.247	20	.556736	.5243
50.572	21	.574612	.7481
51.528	21	.587548	.6218
51.842	22	.589984	.6188
53.203	18.5	.623235	.9057
53.627	20	.628021	.80178
54.679	20	.635383	.8534
55.196	20	.649020	.9674
55.653	22	.657830	.9559
56.520	19	.679423	.9967
56.795	19.5	.690584	3.0252
57.275	25	.689207	.1324
57.967	20	.701399	.0376

Water + Cupric sulfate ($CuSO_4$)

Heterogeneous equilibria .

Tammann, 1885

t	p 0%	p 17%	p 21.85%
46.13	76.2	73.1	46.13
53.90	112.0	107.6	53.90
55.83	122.9	117.7	55.83
63.65	176.6	170.0	63.65
63.70	177.0	170.8	63.70
67.22	207.1	199.6	67.22
70.06	234.4	227.6	70.06
73.85	275.6	265.9	73.85
77.70	323.6	313.3	77.70
80.39	361.0	348.0	80.39
83.49	408.8	394.8	83.49
86.69	463.2	449.2	86.69
89.67	519.5	503.1	89.67
92.28	573.1	555.2	92.28
94.74	627.9	607.4	94.74
97.02	682.6	661.3	97.02
99.99	759.5	736.6	99.99

%	p
100°	
9.80	751.7
19.28	741.0
27.19	726.8
27.60	722.2
35.95	700.2

Emden, 1887

t	p	t	p	t	p
5.80%		10.64%		14.38%	
20.33	17.65	20.26	17.4	19.64	16.6
25.38	23.9	24.55	22.6	26.28	24.9
30.56	32.2	31.22	33.55	30.02	31.05
31.06	33.3	34.96	41.0	35.47	41.95
34.69	40.6	40.31	55.1	39.92	53.35
35.98	43.8	44.36	67.6	44.81	69.0
40.71	56.2	51.07	95.55	49.82	88.95
45.75	73.25	54.05	110.05	55.72	118.5
50.40	93.0	60.90	152.6	60.45	148.2
54.69	114.4	64.55	179.25	65.04	182.1
61.17	154.85	70.05	229.4	69.81	224.95
64.64	181.95	75.64	290.7	75.91	292.1
70.68	236.3	80.03	347.7	79.52	338.6
75.90	296.2	85.58	435.75	85.78	436.2
80.10	351.6	94.98	621.0	90.71	527.5
84.66	420.8			95.67	635.1
89.93	517.8				
95.07	628.3				

Schüller, 1889-90

t	p 0%	p 4.75%	p 11.31%	p 16.67%
15.00	12.700	12.577	12.479	12.339
		12.614	12.485	12.357
19.20	16.555	16.452	16.298	16.129
		16.449	16.291	16.132
21.30	18.844	18.718	18.572	18.402
		18.727	18.553	18.375
23.00	21.665	21.524	21.344	21.144
		21.535	21.339	21.143
26.30	26.202	26.053	25.768	25.584
		26.050	25.822	25.588
27.50	27.303	27.129	26.913	26.664
		27.146	26.910	26.668
30.50	32.534	32.326	32.089	31.897
		32.353	32.082	31.810
33.20	37.842	37.640	37.327	36.951
		37.638	37.332	37.028
33.95	39.457	39.250	38.947	38.620
		39.245	38.932	38.614
39.00	52.039	51.781	51.442	50.995
		51.775	51.380	50.984
40.60	56.708	56.425	55.997	55.562
		56.425	56.000	55.576
51.00	96.661	96.205	95.469	94.813
		96.218	95.454	94.890
60.05	149.143	148.460	147.503	146.227
		148.492	147.515	146.539
70.40	237.213	236.195	234.660	232.997
		236.195	234.668	233.140
73.20	267.442	-	264.641	262.740
		-	264.570	262.845
74.20	279.013	277.797	-	-
		277.801	-	-
75.00	288.517	287.305	285.451	283.494
		287.272	285.404	282.537
75.35	292.820	-	289.582	-
		-	289.661	-
81.00	369.287	367.667	365.283	362.953
		367.663	365.229	362.791
84.50	424.669	424.411	419.785	417.164
		424.769	419.920	417.071

Smits, 1899-1900

m	p	m	p
0°			
0	4.579	0.49378	4.546
.02348	.578	0.99612	.511
.09860	.574	1.2162	.482
.24519	.563		

Pareau, 1877

t	p diss.	t	p diss.
(5+1)			
19.8	5.3	46.1	36.5
26.9	9.1	48.3	39.2
34.7	15.7	50.1	46.2
42.1	28.9	52.9	52.3
44.1	29.8	54.1	58.2

Lescoeur, 1886 and 1890

t	p dissoc. (1+1)	p dissoc. (3+1)	p dissoc. (5+1)
10	-	-	2.8
15	-	-	4
20	-	-	6
25	-	-	8.5
30	-	5	12.5
35	-	7.5	17
40	-	11	23
60	-	45	72
80	-	168	263
100	-	525	688
163	11	-	-
186.5	44	-	-
206	143	-	-
220	666	-	-

Frowein, 1887

t	p diss.	t	p diss.
(5+1)			
13.95	2.993	34.75	15.307
20.46	5.056	39.55	21.452
26.30	8.074	39.70	21.726
30.20	10.987		

Cumming, 1909

t	p dissoc.
(5+1)-(3+1)	
20.1	5.2
25.9	7.8

Ishikawa and Murooka, 1933

t	p dissoc.	t	p dissoc.
(5+1)			
20	16.99	35	40.63
25	22.99	40	53.37
30	30.80		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
25°			
0.1	0.561	0.7	0.458
0.2	0.515	0.8	0.457
0.3	0.494	0.9	0.458
0.4	0.478	1.0	0.461
0.5	0.469	1.2	0.473
0.6	0.462	1.4	0.491

Gerlach, 1886

%	b.t.	%	b.t.
0	100	38.65	102.5
17.55	100.5	40.83	103
26.96	101	42.83	103.5
32.43	101.5	44.47	104
35.98	102	45.11	104.2

Kahlenberg, 1901

%	b.t.	%	b.t.
760 mm			
3.24	100.09	28.35	101.22
7.24	100.19	30.56	101.44
10.87	100.19	32.72	101.73
13.76	100.38	34.22	101.96
18.64	100.58	36.29	102.33
21.55	100.72	39.17	102.96
24.44	100.89	40.82	103.44
26.71	101.07	42.46	103.85

Poggiale, 1843

%	f.t.	%	f.t.
15.39	0	27.97	60
17.30	10	31.06	70
19.06	20	34.70	80
20.84	30	39.11	90
23.25	40	42.97	100
25.45	50		

Tobler, 1855

%	f.t.
14.5	0
19.5	20
22.2	35
26.5	54

Mulder, 1864 and 1866

%	f.t.	%	f.t.
13.22	0	27.33	60
15.90	8.25	29.18	65
16.81	16	34.60	79
17.42	19	42.00	99
20.82	32	43.80	104
21.38	34.5		

Rüdorff, 1872

%	f.t.	%	f.t.
5.81	-0.65	10.65	-1.30
6.82	-0.85	12.36	-1.55
8.34	-1.00	12.78	-1.60
8.82	-1.05	13.98	-1.80

de Coppet, 1872

%	f.t.	%	f.t.
5.81	-0.72	16.60	-2.45
6.82	-1.05	19.28	-2.8
10.65	-1.4	21.33	-3.9
14.77	-2.1		

Tilden and Shenstone, 1885

%	f.t.	%	f.t.
42.70	188	46.10	135
45.05	157	47.40	120
45.77	140		

Etard, 1894

%	f.t.	%	f.t.
12.1	-1	38.8	88
14.1	+7	38.9	89
14.5	9	41.8	94
16.9	18	41.9	96
17.3	20	42.0	97
21.3	35	43.6	100
21.8	39	43.8	108
23.9	45	43.4	110
26.8	54	43.9	116
28.8	61	44.8	120
29.1	63	44.8	132
30.0	65	44.7	133
31.6	70	45.0	143
32.6	71.7	44.2	160
34.5	76	44.5	165
36.6	80	42.9	179
37.8	86	42.2	189

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.072	-0.172	0.595	-0.866
.144	.312	0.890	-1.275
.476	.714	1.190	-1.740

Cohen, Chattaway and Tombrock, 1907

%	f.t.	%	f.t.
	(5+1)		
12.40	0	17.20	20.0
15.02	10.0	18.23	25.0
16.15	15.0		

Schreinemakers, 1910-11

mol %	f.t.
(5+1)	
2.921	35
.454	25
.11	15

Lattey, 1923

%	f.t.	%	f.t.
15.77	13.8	25.54	50.8
18.64	25.0	27.33	55.5
18.73	25.2	28.42	58.8
18.84	25.5		

Ishikawa, 1923

%	f.t.	%	f.t.
17.298	20.0	19.781	30.0
17.886	22.5	20.427	32.5
18.515	25.0	21.085	35.0
19.128	27.5		

Agde and Barkholt, 1926

%	f.t.	%	f.t.
42.19	100	22.26	40
40.48	96	19.55	30
39.23	93	17.02	20
34.89	80	14.98	10
32.02	70	12.90	+0.5
28.39	60	12.70	-1.4
25.20	50		

Properties of phases.

Wiedemann, 1856

%	d	%	d
15°			
0	0.9991	5.630	1.0609
1.949	1.0227	7.374	.0800
3.8233	.0415	9.055	.0993
4.743	.0492	10.682	.1182

Schiff, 1858-59

%	d	%	d
18°			
0	0.9986	9.59	1.0978
0.64	1.0049	10.23	.1048
1.28	.0112	10.87	.1120
1.92	.0176	11.51	.1193
2.56	.0240	12.15	.1265
3.20	.0305	12.78	.1338
3.83	.0370	13.42	.1411
4.23	.0436	14.06	.1485
5.10	.0501	14.70	.1569
5.75	.0567	15.94	.1643
6.39	.0634	15.98	.1722
7.03	.0701	16.62	.1801
7.67	.0770	17.26	.1882
8.31	.0839	17.90	.1963
8.95	.0908	18.54	.2047
		19.17	.2129

Fouqué, 1867

%	d 0°	d t
1.53	1.0109	1.0104 (12°)
3.75	.0339	.0338 (6°)
16.13	.1376	.1352 (13°)

Ewing and Mc Gregor, 1872-73

%	d	%	d
10°			
0	0.9997	15.00	1.1174
2.19	1.0167	17.49	.1386
2.90	.0216	18.00	.1432
4.28	.0318	22.50	.1823
8.18	.0622	25.02	.2051
11.25	.0858	27.29	.1679

t	d	t	d
0.18%		0.57%	
14.63	1.0031	13.72	1.0096
18.90	.0024	18.50	.0087
23.07	.0014	21.34	.0080
27.54	.0003	26.02	.0069
31.81	0.9991	30.89	.0054
1.12%		2.28%	
12.81	1.0158	13.92	1.0254
16.67	.0150	16.57	.0250
21.14	.0140	24.29	.0230
24.80	.0132	26.42	.0224
28.96	.0120	30.49	.0212
34.15	.0104	36.08	.0192
3.72%		4.42%	
14.94	1.0408	15.04	1.0449
18.09	.0403	19.11	.0445
22.87	.0392	19.82	.0438
27.54	.0379	23.17	.0430
32.32	.0367	27.74	.0418
		28.86	.0415
		34.55	.0391
		39.13	.0374
		44.31	.0352
10.73%		12.73%	
13.52	1.1135	15.55	1.1357
18.70	.1118	20.83	.1337
21.34	.1107	28.15	.1311
24.70	.1095	34.35	.1284
28.76	.1080	46.04	.1230
32.22	.1065		
39.53	.1037		
53.76	.0955		
16.38%		16.69%	
19.11	1.1784	14.13	1.1840
24.80	.1760	19.31	.1823
29.78	.1734	25.71	.1795
38.62	.1696	30.39	.1771
44.00	.1674	37.40	.1741

Kohlrausch, 1879

%	d
18°	
5	1.0513
10	.1073
15	.1675
17.5	.2003

Le Blanc, 1889

%	d
20°	
0	0.99823
5.58	1.05687
15.05	.16728
16.79	.18914

Schüller, 1889-90

t	d 4.75%	d 11.31%	d 16.67%
21.0	1.04719	1.10452	1.18976
21.8	.04715	.10443	.18962
40.2	.04661	.10313	.18758
59.5	.04614	.10226	.18637
69.6	.04598	.10201	.18602

Pann, 1901

%	d 10°	d 18°	d 30°
0	0.999727	0.998622	0.995673
6.43	1.07103	1.06922	1.06490
12.33	.13762	.13484	.13151
14.67	.16765	.16246	-
25.40	-	-	.22348

Brümmer, 1902

15°

%	d	%	d
0	0.9993	10.091	1.1084
3.445	1.0353	13.370	.1478
6.764	1.0710	15.391	.1726

Vaillant, 1905

15°

N	d	N	d
0.7856	1.1007	2.0113	1.2218
1.3425	.1494	.3510	.2588
1.6499	.1829	.7213	.3137

Chéneveau, 1907

% (15°)	d (15°)	% (18°)	d (18°)
0	0.9991	0	0.99860
4.73	1.0484	2.35	1.0248
9.06	.0967	4.62	.0496
13.02	.1440	5.70	.0611
16.67	.1897	6.77	.0735
		8.83	.0973
		10.79	.1210

Varga, 1911

18°

N	d	N	d
0.04450	1.002353	1.13331	1.087847
.09415	.006327	.30157	.100848
.22489	.016930	.43247	.110918
.34398	.026411	.77673	.137029
.45651	.035237	.93247	.148688
.62691	.048706	2.27669	.174188
.79942	.062132	.45362	.187519

Keitanpää and Rantaneh, 1911-12

N	d 20°	d 30°	d 40°	d 50°
0.1	1.0087	1.0084	1.0083	1.0082
.2	.0161	.0157	.0156	.0154
.5	.0396	.0391	.0387	.0385
.8	.0619	.0613	.0610	.0610
1.0	.0777	.0767	.0761	.0759
2.0	.1464	.1451	.1446	.1442

Herz, 1914

25°

N	d	N	d
0	0.9971	1.704	1.1281
0.568	1.0407	2.272	1.1708
1.136	1.0848		

Irueste, 1915

%	d	$\alpha.10^6$	$\beta.10^7$
0	0.99913	-	-
1.02	1.00974	155	51
2.73	.02852	184	45
5.92	.06210	225	40
12.31	.13546	285	30
15.02	.16948	315	27

$d^t = d^{15}(1-\alpha(t-15)-\beta(t-15)^2)$

Herz, 1917

25°

N	d	N	d
0	0.9971	1.704	1.1280
0.568	1.0411	2.272	1.1709
1.136	1.0853		

Suominen, 1922

t	d	t	d	t	d
2N		1.5N		1N	
0.0	1.1503	0.0	1.1159	0.0	1.0816
10.7	.1471	12.7	.1136	9.6	.0804
20.7	.1449	20.1	.1113	21.7	.0770
32.6	.1395	32.3	.1064	30.0	.0743
40.0	.1369	41.0	.1024	42.1	.0682
50.1	.1303	51.0	.0974	50.7	.0632
59.3	.1245	61.0	.0915	63.1	.0585
71.4	.1169	75.4	.0821	73.0	.0523
79.0	.1112	87.5	.0745	85.2	.0442
85.0	.1068				
0.5N		0.25N		0.1N	
0.0	1.0408	0.0	1.0221	0.0	1.0100
9.5	.0401	10.0	.0213	9.3	.0096
20.0	.0380	22.2	.0194	21.2	.0076
30.6	.0342	33.0	.0157	32.2	.0046
39.3	.0310	41.2	.0126	43.4	.0007
52.2	.0247	51.6	.0077	50.4	0.9973
67.5	.0175	64.0	.0014	59.8	.9924
71.0	.0144	71.9	0.9969	73.4	.9854
		81.4	0.9881	83.2	.9787

Rakshit, 1925

%	d	%	d
0.64	1.00492	6.40	1.06150
3.20	1.03027	19.20	1.18221

Agde and Barkholt, 1926

%	t	d	%	t	d
12.70	- 1.4	1.146	28.39	60	1.343
12.90	+ 0.5	.149	32.02	70	.398
14.98	+10	.176	34.89	80	.433
17.02	20	.198	39.23	93	.504
19.55	30	.232	40.48	96	.521
22.26	40	.262	42.19	100	.550
25.20	50	.299			

Flöttmann, 1928

%	t	d
16.12	15	1.1839
17.30	20	1.1965
18.49	25	1.2111

Gibson, 1934

mol %	d
	25°
0	0.9970
10	1.1055
15	1.1645

Pearce and Pumplin, 1937

M	d	M	d
	25°		
0.0	0.997074	0.8	1.121023
.1	1.013235	1.0	.150589
.2	.029136	1.1964	.179083
.4	.060390	1.4182	.210706
.6	.090990	satd.	

Berthier, Courty and Gauthier, 1951

%	t	d	%	t	d
18.11	22.5	1.2062	7.01	23	1.0720
17.44	24.5	.1979	5.93	24	.0601
16.47	20	.1895	5.40	24	.0547
16.44	17.5	.1856	4.47	22	.0453
15.37	28	.1687	4.44	23	.0442
14.86	16	.1677	3.38	23	.0331
13.55	18	.1500	2.38	23	.0223
12.14	19	.1342	1.46	24	.0118
11.13	20	.1203	1.11	23	.0091
10.41	20.5	.1106	0.849	22	.0071
9.16	21	.0965	.568	23	.0043
8.42	22	.0890	.450	24	.0024
7.64	22.5	.0810	.251	22.5	.0003

Gibson, 1934

mol%	π(1-1000 bars)
25°	
0	39.46
10	30.66
15	26.75

Baranowski and Demichowiez, 1953 (fig.)

Electrothermodiffusion

N	Nl/Nu*	N	Nl/Nu*
0.02	1.02	0.20	2.18
.03	.30	.40	2.00
.04	.60	.60	1.89
.07	.90	.80	1.84
.10	2.10	1.00	1.86

* Nl and Nu are concentrations at lower and upper ends of a capillary of 1 mm external diameter and 10 cm length, for a current of 35 w during 2 hours.

Wagner, 1883

%	η (water°=1) 15°	25°	35°	45°
6.79	0.7963	0.6181	0.4978	0.4140
12.57	0.9818	0.7398	0.5974	0.5200
17.49	1.2452	0.9685	0.7588	0.6180

Keitanpää and Rantaneh, 1911-12

N	η (water=1) 20°	30°	40°	50°
0.1	1.0333	1.0321	1.0290	1.0255
.2	.0586	.0534	.0502	.0469
.5	.1690	.1631	.1536	.1445
.8	.2539	.2442	.2312	.2258
1.0	.3211	.3167	.2946	.2915
2.0	.7651	.7434	.7136	.6750

Herz, 1914 and 1917

N	η
25°	
0	895
0.568	1074
1.136	1293
1.704	1541
2.272	1836

Wiedemann, 1856

%	η	%	η
	15°		
0	1131	5.630	1478
1.949	1192	7.374	1639
3.8233	1361	9.055	1800
4.743	1418	10.682	1973

Volkmann, 1882

d	σ
15-16°	
0.999	73.2
1.0276	74.2
.0611	74.5
.1189	75.0
.1775	75.8

Kasankin, 1891

d 16-17°	capillary rise	d 32-34°	capillary rise
1.1864	23.23	1.2562	21.73

Pann, 1901

%	σ 10°	30°
0	74.10	71.02
6.43	74.73	71.75
12.33	75.54	72.60
14.67	75.99	-
25.4	-	74.00

Brümmer, 1902

%	σ	%	σ
	15°		
0	74.92	10.091	70.77 (?)
3.445	75.06	13.370	75.86
6.764	74.80	15.391	81.15

Forch, 1905

N	t	σ	N	t	σ
0	15	73.26	1.550	17.7	77.94
.350	16.4	76.89	1.860	"	78.19
.558	16.2	77.14	2.325	"	78.61
1.162	16.2	77.74			

Hoeltzenbein, 1924

%	t	diffusion ratio
3.85	10.6	0.35031
7.4	10.6	.27338
13.8	11.2	.27265

Optical and electrical properties.

Le Blanc, 1889

%	n_D
20°	
0	1.33325
5.58	1.34376
15.05	1.36260
16.79	1.36613

Walter, 1889

%	n_D
15°	
0	1.3334
2.45	.3405
6.40	.3480
10.8	.3563
16.8	.3681

Chèneveau, 1907

%	n_D	%	n_D
15°			
0	1.3335	13.02	1.3585
4.73	1.3420	16.67	1.3662
9.06	1.3503		
18°			
0	1.3331	6.77	1.3462
2.35	1.3379	8.83	1.3502
4.62	1.3421	10.79	1.3541

Flöttmann, 1928

%	t	n_D
16.12	15	1.36544
17.30	20	.36727
18.49	25	.36927

Ewing and Mc Gregor, 1872-73

%	κ	%	κ
10°			
2.19	57.22	17.49	266.0
2.90	69.85	18.00	275.8
4.28	95.00	22.50	307.4
8.18	159.5	25.02	321.7
11.25	198.9	27.29	296.4
15.00	246.9		

Kohlrausch, 1879

%	κ	τ
18°		
2.5	108	0.0214
5	188	.0217
10	319	.0219
15	420	.0232
17.5	457	.0237

Bachofner, 1904

t	κ 0.5N	1N	1.5N	2N	2.5N
10	116	196	263	310	350
15	140	236	310	369	415
20	164	273	357	426	479
25	186	310	403	484	542
30	207	344	449	540	604
35	227	377	493	593	665
40	248	408	534	642	724
45	266	439	575	691	780
50	283	467	612	735	834
55	298	494	646	777	882
60	312	519	680	816	929
65	324	541	707	850	972
70	334	560	732	881	1009
75	343	578	755	908	1044
80	352	594	774	934	1076

M	c	κ 20°	40°	60°	80°
0.1	2.497	90	130	160	180
.2	4.995	140	210	262	295
.25	6.243	164	248	312	352
.3	7.492	188	282	358	407
.4	9.990	233	348	444	507
.5	12.487	273	408	519	594
.6	14.984	310	461	587	669
.7	17.482	342	510	650	740
.75	18.729	357	534	680	774
.8	19.979	372	557	708	807
.9	22.476	400	601	764	872
1.0	24.974	426	642	816	934
.1	27.471	450	678	864	994
.2	29.966	470	710	908	1050
.25	31.255	479	724	929	1076

Cheneveau, 1910

%	χ
23°	
5.66	-0.397
11.27	0
16.9	+0.416
22.4	+0.800

Cabrera and Moles, 1915

%	χ^{mol}_{Cu}	%	χ^{mol}_{Cu}
20°			
1.503	1621	11.24	1607
3.045	1632	14.60	1585
5.36	1624	15.77	1596
7.72	1606		

Berthier, Courty and Gauthier, 1951

%	t	χ	%	t	χ
18.11	22.5	+1.256	7.01	23	+0.073
17.44	24.5	.180	5.93	24	-0.057
16.47	20	.140	5.40	24	.114
16.44	17.5	.090	4.47	22	.215
15.37	28	0.943	4.44	23	.209
14.86	16	.948	3.38	23	.323
13.55	18	.816	2.38	23	.435
12.14	19	.657	1.46	24	.556
11.13	20	.494	1.11	23	.587
10.41	20.5	.421	0.849	22	.591
9.16	21	.290	.568	23	.632
8.42	22	.218	.450	24	.651
7.64	22.5	+0.127	.251	22.5	-0.682

Heat constants .

Gray, 1879 - 1880

d	U
at room temperature	
1.184	0.8354
.142	.8788
.109	.8998
.0871	.9193

Holsboer, 1902

mol%	U	mol%	U
0.2	0.972	3.9	0.696
.5	.940	4.75	.656
.65	.925	5.55	.638
1.0	.893	6.25	.615
2.0	.813	6.5	.610

Vaillant, 1905

N	U	N	U
15°			
0.7856	0.9325	2.0113	0.8478
1.3425	.8893	.3510	.8288
1.6499	.8709	.7213	.8094

Kapustinskii, Yakuchevskii and Drakin, 1953

%	U	%	U
0	0.9980	8.27	0.9072
3.91	.9512	10.36	.8881
6.75	.9211	12.34	.8718

Scholz, 1892

M	Q diss (cal/g)
0.125	13.60
.250	13.38
.5	12.98
1.0	12.64
2.5	11.48

Holsboer, 1902

mol%	Q diss (cal/mole(8+3))				
	15°	10°	5°	20°	25°
6.15	870	966	1061	775	679
5.47	1258	1332	1405	1185	1111
4.30	1633	1645	1657	1621	1609
3.01	1918	1876	1835	1959	2001
1.88	2118	2065	2013	2170	2223
0.97	2288	2203	2118	2373	2458
0.49	2418	2306	2194	2530	2642
0.25	2530	2303	2075	2758	2085
0.20	2568	-	-	-	-

Donnan and Hope, 1910

mol% (initial)	Q dil*	mol% (initial)	Q dil*
16.67	-2750	25.99	+3720
19.35	+65	36.01	7910
23.17	2123	36.63	7840
30.98	5425	41.36	9510
49.30	9229	55.25	11000
100	15820	60.93	11600
		71.78	13100
17.21	-2600	74.25	13300
18.07	-1860	78.49	13500
19.05	-660	84.03	14300
20.83	+570	88.82	14807
		93.78	15100

* by mole salt in 400 mole aq.

Weber, 1880

d	heat conductivity	U
	15°	
1.000	0.0745	1.000
1.160	.0710	0.848

Water + Cupric dithionate (CuO_6S_2)

De Baat, 1926

%	f.t.
43.82	0
44.91	20
45.51	30

Water + Cupric sodium sulfate ($CuNa_2O_8S_2$)

Koppel, 1902

wt%	mol%	f.t.
27.68	2.24	17.7
27.44	2.25	19.5
27.12	2.22	23
26.44	2.14	30
25.99	2.10	40.15

t dissoc.= 16.7°

Water + Cupric diammonium chloride ($CuN_2H_8Cl_4$)

Rüdorff, 1862

%	f.t.	%	f.t.
	(2+1)		
5.12	-2.65	12.48	-7.3
5.64	-2.9	18.24	-11.55
6.68	-3.5	19.40	-12.60

Jones, 1904; Jones and Bassett, 1905

M	f.t.	d	molecular conductivity
			0°
0.05	-0.600	1.007900	248.6
.10	-1.214	.015500	235.1
.20	-2.515	.031700	205.6
.30	-3.964	.048300	197.6
.40	-5.500	.063500	187.4
.60	-8.700	.092800	162.9
.80	-12.000	.122700	146.1
1.00	-	-	130.4

Meerburg, 1905

%	f.t.	%	f.t.
3.87	-1.5	21.16	-5
5.88	-2.48	22.02	0
8.78	-3.95	24.26	+12
9.97	-4.60	25.95	20
13.12	-6.40	27.70	30
15.84	-8.04	30.47	40
17.64	-9.24	33.24	50
20.12	-10.80	36.13	60
20.30	-11.0 E	39.35	70
20.46	-10 (2+1)	43.36	80

Water + Cupric diammonium sulfate ($CuN_2H_8O_8S_2$)

Beetz, 1879

d^{20}	heat conductivity	t
1.086	415	8-14°
"	659	28-36°
water	413	8-14°
"	662	28-36°

Water + Cupric acetate ($CuC_4H_6O_4$)

Heydweiller, 1912

N	d	κ
	18°	
0.1003	1.00411	23.04
.1994	.00928	34.52
.4978	.02471	53.94
.9403	.03694	62.34
1.796	.03973	63.77

Water + Cupric sulfamate ($CuH_4N_2O_6S_2$)

Baker, 1949

M	d	M	d
	25°		
0.08283	1.01340	1.2199	1.22814
.13250	.02316	.5304	.28484
.54394	.10238	.9495	.36007
.97910	.18428	2.4140	.44351

42.77% f.t.= 25°

Water + Cupric benzene sulfonate ($CuC_{12}H_{10}O_6S_2$)

Ephraim and Seger, 1925

c	f.t.	c	f.t.
32.325	18	48.891	50
39.255	34	55.920	64.5
49.701	49.5	66.240	80.5

Water + Manganese chloride ($MnCl_2$)
Heterogeneous equilibria.

Tammann, 1885

%	p	%	p
	100°		
8.12	738.1	30.77	612.1
13.68	716.0	34.04	588.5
22.50	670.2	39.77	541.0

Dawson and Williams, 1899

t	p sat.sol.	p diss. (4+1)	p diss. (2+1)
15.5	7.85	3.50	-
20.0	9.45	4.06	-
30.0	17.02	9.31	2.9
40.0	28.67	19.28	5.9
50.0	46.29	41.72	-
57.8	62.9	62.83	13.8
60.0	64.8	-	26.0
70.5	110.25	-	38.4
80.5	172.1	-	73.2
90.0	262.0	-	-

Perreu, 1930

%	p	%	p
	20°		
27.26	13.85	38.16	11.44
31.80	12.91	40.05	10.98
35.33	12.14	41.67	10.58

Robinson and Stokes, 1940

Isopiestic solutions

m_1	m_2	m_1	m_2
	25°		
0.1148	0.1596	0.1463	0.2044
0.4537	0.6801	0.6091	0.9450
1.329	2.396	1.539	2.868
2.306	4.662	2.353	4.782
0.2276	0.3248	0.2937	0.4251
0.9826	1.660	1.259	2.236
1.601	3.000	1.966	3.860

$_1$= manganese chloride

$_2$= potassium chloride

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.853	1.6	1.173
0.2	0.859	1.8	1.221
0.3	0.872	2.0	1.264
0.4	0.889	2.5	1.366
0.5	0.908	3.0	1.454
0.6	0.929	3.5	1.528
0.7	0.950	4.0	1.584
0.8	0.971	4.5	1.634
0.9	0.995	5.0	1.671
1.0	1.022	6.0	1.735
1.2	1.072	7.0	1.798
1.4	1.124	8.0	1.861

Lescoeur, 1890

t	p dissoc.	t	p dissoc.
(4+1)		(2+1)	
20	3.8	100	96
30	18		
40	31		
50	46		
60	78		
70	125		
80	182		

Rüdorff, 1862

%	f.t.	%	f.t.
2.45	-1.0	10.37	-5.55
2.57	-1.05	15.95	-10.85
5.59	-2.45	16.87	-12.05

Dupuy, 1884

%	f.t.	%	f.t.
42.54	15	52.53	60
47.29	30	58.84	100

Etard, 1894

%	f.t.	%	f.t.
34.7	-22	48.2	55
37.8	-5	50.0	57
40.4	+7	51.0	80
41.2	17	53.7	100
42.3	19	54.7	140
44.4	35		

Dawson and Williams, 1899

%	f.t.	%	f.t.
43.56	25	51.31	57.65
44.66	30	52.06	60
46.98	40	52.52	70
49.53	50	52.99	80

tr.t.: (4+1)-(2+1) 59° (2+1)-anh. 198°

Jones, 1904; Jones and Bassett, 1905; Jones and Getman, 1904

M	f.t.	M	f.t.
0.059	-0.255	1.061	-5.965
.106	-0.508	1.500	-11.100
.133	-0.639	2.000	-16.500
.266	-1.259	2.500	-24.000
.400	-2.004	3.000	-31.000
.532	-2.790	3.500	-40.000
.796	-4.247	4.000	-48.500
.902	-4.825		

Kuznetsov, 1909

(4+1) I f.t.= 58.2° (4+1) II f.t.= 40°

Benrath, 1934

%	f.t.	%	f.t.
38.86	0 (4+1)	51.60	60 tr.t.
42.36	20	52.03	80 (2+1)
43.67	27	52.37	90
45.46	35	53.40	99

Benrath, 1941

%	f.t.	%	f.t.	%	f.t.
54.5	113	64.5	217	79.5	333
55.0	121	66.9	249	81.5	345
56.7	146	69.6	272	82.5	350
57.9	155	71.3	286	83.5	353
58.6	176	75.1	313	85.5	392
61.3	189	76.5	321	85.5	430
63.7	198	78.5	328	100	650

E : 30.5 % -26.5°

tr.t. (6+1) - (4+1) 38.5 % - 2°
(4+1) - (2+1) 51.4 % +58.1°
(2+1) - (1+1) 63.7 % 198°
(1+1) - anh. 85.0 % 362°

Properties of phases.

Long, 1880

%	d	%	d
	15°		
5	1.0456	20	1.1900
10	.0895	25	.2472
15	.1378	28	.2828

Blümcke, 1884

%	t	d
10.07	15	1.070
20.02	16	.151
30.05	15	.241
40.24	15.5	.345
49.97	16	.460

Jahn, 1891

c	d
0	0.9982
11.060	1.0874
25.155	1.1964

Sentis, 1897

mol%	t	d	mol%	t	d
1	16.9	1.0551	5	18.0	1.2557
2	17.0	.1067	6	17.85	.3013
3	17.4	.1568	7	18.1	.3472
4	17.9	.2076	8	18.4	.3932

Dawson and Williams, 1899

%	t	d	%	t	d
43.56	25	1.4991	51.31	57.65	1.6097
44.66	30	.5049	52.06	60	.6108
46.98	40	.5348	52.52	70	.6134
49.53	50	.5744			

Jones, 1904; Jones and Bassett, 1905; Jones and Getman, 1904

M	d	M	d
	0°		
0.133	1.010716	1.500	1.148996
.266	.024500	2.000	.195664
.400	.037876	2.500	.241484
.532	.051064	3.000	.282164
.796	.077808	3.500	.330580
.902	.087480	4.000	.383708
1.061	.106532		

Herz, 1914 and 1917

N	d
25°	
0	0.9971
1.640	1.0800
3.280	.1607
4.920	.2396
6.560	.3175

Cabrera, Moles and Marquina, 1918

%	d	$\alpha.10^5$	%	d	$\alpha.10^5$
0	0.9982	-	8.222	1.0724	19
1.566	1.0120	15	15.66	.1437	22
3.489	.0286	17	27.70	.2784	27
5.374	.0454	18	38.42	.4257	31

$d^t = d^{20}\ (1-\alpha(t-20))$

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.02589	4	1.1975
1	.0512	5	1.2448
2	.1010	6	1.2917
3	.1496		

Duperier, 1924; Cabrera and Duperier, 1925

%	d
20°	
3.028	1.0250
15.035	.1385
37.90	.4175

Guillaume, 1946

%	d
20°	
5.5	1.049
19.7	1.1862

Wagner, 1883

%	η (water°=1) 15°	25°	35°	45°
8.007	0.9281	0.7112	0.5746	0.4811
15.65	1.3090	1.0420	0.8401	0.6869
30.33	2.5630	1.9320	1.5500	1.2370
40.132	5.3730	3.9340	2.0040	2.4650

Herz, 1914 and 1917

N	η	N	η
25°			
0	895	4.920	2255
1.640	1237	6.560	3294
3.280	1674		

Sentis, 1897

%	t	σ	%	t	σ
0	25.1	72.3	4	17.9	80.3
0	13.5	74.0	5	18.0	82.2
1	16.9	75.2	6	17.85	83.7
2	17.0	76.8	7	18.1	85.3
3	17.4	78.3	8	18.4	87.5

Jahn, 1891

c	n H_α	n D	n H_β
0	1.3315	1.3332	1.3375
11.060	1.3536	1.3556	1.3603
25.155	1.3800	1.3823	1.3876

Jones, 1904; Jones and Bassett, 1905; Jones and Getman, 1904

M	n_D	M	n_D
0°			
0.059	1.32656	0.532	1.33768
.106	.32766	.796	.34363
.133	.32829	.902	.34588
.266	.33150		

Heydweiller and Grube, 1916

w.l.(in Å)	D n_D . 10^5 (sol.-aq.) 0.4995 N	1.015 N	2.024 N	4.042 N
18°				
2314	931	1869	3654	7005
2574	840	1683	3295	6359
2749	801	1608	3145	6066
2981	766	1537	3005	5800
3256	738	1479	2894	5583
3405	727	1457	2848	5493
3612	715	1431	2797	5398
4679	683	1345	2647	5106

Jahn, 1891

c	$(\alpha)_{magn.}$
20°	
0	1
11.060	1.3227
25.155	1.2381

Guillaume, 1946

%	(α)*magn.10^6	n 5780 Å
20°		
5.5	4.083	1.3467
19.7	4.206	1.3816

*in radians, gauss, centim.

Mc Clung and Mc Intosh, 1902

d	% of X-ray absorption
at room temperature	
1.194	96.9
.097	89.6
.048	83.0
.023	70.7
.010	63.5
.000	61.2

Long, 1880

%	ϰ	τ	%	ϰ	τ
		18°			
5	500	0.0210	20	1078	0.0206
10	802	.0206	25	1036	.0203
15	799	.0202	28	965	.0208

Jones, 1904; Jones and Bassett, 1905; Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.059	88.67	1.500	43.35
.106	81.81	2.000	34.82
.133	79.85	2.500	26.38
.266	73.49	3.000	19.67
.532	65.16	3.500	14.04
.796	57.32	4.000	9.82
1.061	51.44		

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	66.4	4	28.17
1	57.4	5	21.8
2	45.27	6	16.30
3	35.7		

Cabrera, Moles and Marquina, 1918

%	χ mol	%	χ mol
	18°		
0.582	14893	8.222	14991
1.566	14967	15.660	15000
3.489	14922	27.700	14987
5.374	14879	38.480	14955

Duperier, 1924; Cabrera and Duperier, 1925

%	χ
20°	
3.028	2.9
15.035	17.2
27.90	44.4

Blümcke, 1884

%	U	%	U
	15°		
10.07	0.891	40.24	0.668
20.02	0.810	49.97	0.608
30.05	0.733		

Water + Manganese bromide ($MnBr_2$)

Etard, 1894

%	f.t.	%	f.t.
52.1	-21	68.2	64
56.5	+7	70.1	76
57.0	11	69.7	89
59.1	18	69.2	97
62.7	38	70.1	105
64.2	52		

Kuznetsov, 1909

tr.t. (4+1) II - (2+1) 57-58°

Lescoeur, 1894

t	diss. (1+1) P	sol.sat.
20	-	5
60	-	40
100	-	200
140	56	-
150	99	-
155	122	-
160	156	-

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.04478	72.4
1	.0889	64.7
2	.1769	53.54
3	.2630	43.6
4	.3491	35.64

Water + Manganese iodide (MnI_2)

Kuznetsov, 1909

%	f.t.
63.73	-11.2 (9+1)
64.06	-21.6 (4+1)
69.47	-10.7
71.86	0
72.90	+17.5

Lescoeur, 1894

t	p dissoc. (1+1)	p dissoc. (2+1)	p dissoc. (4+1)	p sat.sol.
20	-	-	-	3.5
50	-	-	5	7
60	-	-	-	13
66	-	19	-	-
80	-	-	34	-
100	-	47	83	42.5
110	-	91	135	104
120	130	112	-	154
130	180	-	-	210
133	198	-	-	-

Water + Manganese nitrate (MnN_2O_6)

Ewing, Glick and Rasmussen, 1942

%	m	p 20°	p 25°	p 30°	p 40°
16.89	1.14	16.66	22.43	30.00	52.03
31.17	2.53	14.90	19.96	26.47	45.94
37.18	3.31	13.07	17.75	23.92	41.49
43.70	4.34	11.51	15.51	20.65	36.28
51.61	5.96	8.68	11.71	15.85	28.04
56.91	7.38	6.83	9.24	12.61	22.38
59.93	8.36	5.61	7.71	10.59	18.87
63.55	9.74	-	5.96	8.09	14.60
66.18	10.94	3.39	4.70	6.47	12.42
68.28	12.03	-	3.69	-	9.39
74.04	15.93	-	1.92	-	5.10
74.19	16.94	-	1.61	-	4.32
77.53	19.28	-	0.98	-	2.69
79.27	21.37	-	0.81	-	2.26
80.32	22.81	-	0.68	-	1.90

t	p	t	p
(6+1)		27.89	6.03
		29.87	6.25
4.01	3.12	30.87	6.43
9.90	4.16	32.88	6.67
15.00	5.36	34.89	6.90
20.04	6.45	35.89	6.70
21.00	6.74	36.90	6.04
21.98	6.88	35.89	4.57
23.32	7.01	34.89	3.77
24.04	6.84	33.89	3.43
24.86	6.64	32.88	3.00
24.87	5.68	30.87	2.16
24.00	4.83	28.88	1.71
21.96	3.61	26.89	1.34
19.98	2.91	24.90	1.10
15.00	1.56	22.91	0.95
(4+1)		(2+1)	
0.00	2.00	26.89	1.20
9.91	2.92	28.88	1.23
15.00	3.69	30.87	1.39
19.94	4.60	32.88	1.48
21.91	5.06	34.89	1.66
22.91	5.21		
23.91	5.39	(6+1)-(4+1)	
24.91	5.67		
24.91	5.60	18.00	3.36
25.41	5.68	20.00	4.02
25.90	5.72	22.00	4.70
26.89	5.93	24.00	5.29

Rüdorff, 1872

%	f.t.
3.45	-0.95
7.67	-2.40
12.51	-4.40
15.68	-6.35
20.64	-9.45

Funk, 1899

%	f.t.	%	f.t.
(6+1)		(3+1)	
29.30	-15.5	65.66	27
32.98	-20	66.99	29
41.70	-30	67.38	30
42.29	-29	71.31	34
43.15	-26	76.82	35.5
44.30	-21		
45.52	-16		
48.88	-5		
50.49	0		
54.50	+11		
57.33	18		
62.37	25.8		
65.92	25.5		

Jones, 1904 and Jones and Bassett, 1905

M	f.t.	M	f.t.
0.27	-1.39	0.09	-0.46
0.54	-2.98	.18	-0.88
1.05	-6.75	.27	-1.43
1.59	-12.00	.54	-3.07
2.61	-27.50	1.05	-6.80
3.15	-38.50	1.59	-11.80
		2.61	-27.00
		3.105	-38.00

Jones and Getman, 1904

mol%	f.t.	mol%	f.t.
0.035	-0.090	1.05	-2.748
.18	-0.410	2.10	-6.310
.35	-0.841	2.80	-9.315
.53	-1.297	3.50	-12.690
.87	-2.184		

Di Capua, 1929

56.81% f.t.= 20°

Ewing and Rasmussen, 1942

%	f.t.		%	f.t.	
52.9	10.3	(6+1)	73.9	34.2	(4+1)
54.8	15.5		75.4	31.3	
56.0	18.1		75.3	31.0	
57.8	21.7		76.1	29.7	
59.8	24.0		76.9	28.4	
61.1	25.0		77.2	26.5	
62.1	25.2		77.9	28.0	(2+1)
62.4	25.3		78.5	30.0	
62.5	25.3		79.2	32.2	
62.7	25.3		79.8	34.3	
63.8	24.9		80.3	35.2	
64.0	24.7	E	80.6	29.0	(1+1)
64.9	27.3		80.9	37.0	
67.2	32.3		81.0	40.1	
68.4	34.5		81.2	45.0	
69.7	36.2		81.6	50.0	
71.3	37.1		82.1	59.9	
71.5	37.0		82.8	68.1	
71.7	36.8		83.3	75.0	
72.2	36.6				

Oudermans, jr., 1868

%	d	%	d
	8°		
0.0	0.9999	23.08	1.2146
0.62	1.0049	23.70	.2214
1.25	.0099	24.32	.2283
1.87	.0150	24.95	.2352
2.49	.0202	25.57	.2421
3.12	.0253	26.19	.2491
3.74	.0305	26.82	.2562
4.37	.0358	27.44	.2633
4.99	.0411	28.07	.2705
5.61	.0464	28.69	.2778
6.24	.0517	29.32	.2851
6.86	.0571	29.94	.2925
7.49	.0626	30.56	.2999
8.19	.0681	31.18	.3074
8.73	.0736	31.80	.3150
9.36	.0792	32.43	.3226
9.98	.0848	33.05	.3303
10.60	.0905	33.67	.3381
11.22	.0963	34.30	.3459
11.84	.1020	34.92	.3538
12.47	.1078	35.55	.3618
13.09	.1137	36.17	.3698
13.71	.1196	36.79	.3779
14.33	.1256	37.42	.3861
14.90	.1316	38.04	.3944
15.59	.1377	38.67	.4027
16.21	.1438	39.29	.4111
16.84	.1500	39.91	.4196
17.46	.1562	40.54	.4281
18.08	.1624	41.16	.4367
18.71	.1688	41.79	.4454
19.38	.1752	42.41	.4542
19.96	.1816	43.03	.4631
20.58	.1881	43.66	.4721
21.20	.1946	44.29	.4811
21.83	.2012		
22.45	.2079		

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	d	M	d
0°			
0.27	1.034668	1.59	1.208012
0.54	.071920	2.61	.336644
1.05	.138448	3.15	.402064

Herz, 1914

N	d
25°	
0	0.9971
1.30	1.0768
2.60	.1558
3.90	.2332
5.20	.3092

Herz, 1917

N	d
25°	
0	0.9971
1.300	1.0767
2.600	.1559
3.900	.2327
5.290	.3092

Cabrera, Moles and Marquina, 1918

%	d	$\alpha.10^5$	%	d	$\alpha.10^5$
20°					
1.884	1.0138	20	18.32	1.1635	33
3.970	.0313	21	38.47	.4023	47
7.238	.0597	24	48.09	.5430	53

$d^t = d^{20}\ (1-\alpha(t-20))$

Heydweiller, 1921

N	d	N	d
18°			
0.5	1.03389	4	1.2618
1	.0672	5	.3243
2	.1332	6	.3857
3	.1981	10	.6215

Cabrera and Duperier, 1925

%	d
20°	
6.420	1.0516
36.90	1.3784

Wagner, 1883

%	η (water°=1) 15°	25°	35°	45°
18.309	0.9596	0.7638	0.6448	0.5560
29.602	1.6750	1.2600	1.0460	0.8865
49.309	3.9685	2.0110	2.2100	1.8880

Herz, 1914 and 1917

N	η
25°	
0	895
1.300	1104
2.600	1358
3.900	1748
5.230	2280

Jones and Getman, 1904

M	n_D	M	n_D
0°			
0.035	1.32554	1.05	1.33768
.18	.32751	2.10	.34974
.35	.32949	2.80	.35734
.53	.33142	3.50	.36459
.87	.33643		

M	λ	M	λ
0°			
0.035	49.18	0.53	39.60
0.18	44.33	0.87	36.80
0.35	41.79		

Jones and Getman, 1904
Jones, 1904 and Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
0°			
0.27	80.62	1.59	48.61
0.54	71.37	2.61	27.37
1.05	58.48	3.15	20.02

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	66.8	4	30.30
1	58.9	5	23.54
2	47.46	6	17.66
3	37.92	10	4.66

Cabrera, Moles and Marquina, 1918

%	χ mol	%	χ mol
	18°		
0.732	14772	18.32	14833
1.884	14734	38.47	14943
3.970	14865	48.09	14941
7.238	14813		

Cabrera and Duperier, 1925

%	χ
20°	
6.420	4.68
36.90	30.30

Ewing, Glick and Rasmussen, 1942

m	Q diss.	m	Q diss.
	25°		
0.03673	-14.40	0.01728	-14.46
.03024	-14.42	.02203	-14.28
.03273	-14.47	.02414	-14.41
.02959	-14.42		

m initial	m final	Q dil.
	25°	
0.815	0.0139	-0.466
3.006	.0371	1.639
5.500	.0493	2.856
10.14	.1015	5.016
11.10	.1194	5.448
14.84	.1301	6.912
17.63	.1329	7.896
21.20	.1034	8.928
24.62	.1033	9.720

Water + Manganese sulfate ($MnSO_4$)

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.587	.2	.485
.2	.538	.4	.504
.3	.516	.6	.527
.4	.501	.8	.556
.5	.490	2.0	.588
.6	.481	2.5	.677
.7	.475	3.0	.782
.8	.472	3.5	.909
.9	.472	4.0	1.048
1.0	.475		

Nicol, 1884

t	p	t	p
	sat.sol.		
65	167.2	85	394.7
75	256.5	95	591.8

Tammann, 1885

t	p 0%	p 24.45%	p 34.49%
41.17	58.9	56.5	51.9
43.32	65.9	63.2	57.3
49.19	88.9	85.3	78.5
52.29	103.6	99.5	91.7
58.73	140.9	135.5	127.8
60.83	155.3	148.5	138.9
66.43	200.0	192.5	180.8
70.70	241.0	231.7	217.2
71.00	244.1	235.0	220.3
74.38	281.9	271.7	255.2
86.16	303.7	293.0	275.4
79.10	342.7	330.9	311.0
82.97	400.5	386.6	367.1
84.78	430.1	415.9	393.3
87.22	473.0	456.9	436.1
90.05	526.9	510.5	485.2
92.39	575.5	557.0	529.1
94.39	620.0	600.6	572.5
97.10	684.6	663.7	639.5
100.22	766.4	743.0	720.1

%	p	%	p
	100°		
4.86	755.8	21.58	741.2
13.86	748.0	28.98	727.9
19.33	744.1		

Bolte, 1912

t	p dissoc.	t	p dissoc.
(7+1)		(2+1)	
14.75	10.55	20.01	3.48
	(5+1)		
9.0	5.63	15.0	9.42
10.0	6.18	15.8	9.95
10.5	6.48	16.0	10.10
11.0	6.80	16.5	10.47
11.6	7.17	17.0	10.81
12.0	7.40	18.0	11.53
12.5	7.72	18.7	12.03
13.0	8.06	19.0	12.21
14.0	8.74	20.0	12.94
14.7	9.20		

Gerlach, 1886

%	b.t.	%	b.t.
0	100	31.60	101.5
14.60	100.5	37.07	102
24.30	101	40.62	102.4

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
3.58	100.12	12.61	100.38
6.65	100.20	16.21	100.53
9.29	100.29	19.49	100.69

Mulder, 1864

%	f.t.	%	f.t.
35.65	0	38.00	63.5
38.80	13.5	38.19	70
43.08	20.5	38.12	78.2
40.80	26	38.19	82.5
41.08	31	37.81	87
42.70	47.25	37.31	92.5
42.99	54.0	35.19	99
42.50	55.3	32.02	102
42.00	57.5	32.29	102.9 b.t.
39.80	60		

Rüdorff, 1872

%	f.t.	%	f.t.
6.10	-0.70	19.20	-3.35
9.77	-1.20	22.40	-4.45
11.20	-1.45	24.12	-5.20
13.00	-1.75	25.28	-6.90
15.50	-2.35	27.67	-7.50

Linebarger, 1893

%	f.t.	%	f.t.	%	f.t.
		(7+1)			
33.38	-10	36.16	7	40.66	35
33.74	- 8	37.46	10	41.29	45
34.00	- 5	39.14	15	42.22	54
34.90	0	39.98	20	42.18	58
35.41	+ 5	40.17	25		
		(6+1)			
35.85	- 4	41.48	9	43.26	35
39.10	0	42.01	15	43.51	45
40.07	+ 3	42.65	20	43.57	54
40.30	+ 5	42.98	25	43.95	58
		(5+1)			
36.73	0	42.91	20	45.84	40
38.43	+ 2.5	44.02	25	46.03	42
39.10	4	44.18	30	46.28	45
40.06	7	44.56	32	46.51	47.7
40.50	10	45.06	34	47.06	53
41.97	15	45.62	37	47.11	54
		(4+1)			
36.66	+ 2.2	41.97	25	45.37	45
38.19	7.3	42.75	30	45.75	48
39.02	11	44.07	35.5	46.26	52
40.16	15	44.33	40	46.86	56
41.16	20				
		(3+1)			
35.35	+ 5	40.07	25	41.82	48
37.72	12	40.26	30	42.13	53
38.80	16	40.59	35	42.26	57
39.36	19	41.39	40		
		(2+1)			
40.79	+35	43.70	42	45.40	50
42.93	40	44.46	45	46.32	55
		(1+1)			
47.36	+48	44.17	78	41.23	106
46.27	53	43.06	90	41.10	115
45.75	65	41.61	100	40.76	117
45.27	72				
		anhydre			
40.18	120	29.17	141	20.94	155
38.71	132	27.97	146	13.90	170

Etard, 1894

%	f.t.	%	f.t.
30.0	- 8	36.4	52
31.0	- 5	41.1	70
34.1	+ 5	36.3 (?)	83
38.3	18	18.4	110
38.2	22	21.5	115
39.1	23	16.7 (?)	123
41.7	32	13.6	130
44.2	45	9.4	140

Cotrell, 1900

%	f.t.	%	f.t.	%	f.t.
		(7+1)			
32.41	-10	38.99	14.3	40.43	35
34.74	0	39.09	18.5	40.55	35.5
36.06	+ 5	39.44	25	40.76	39.9
37.23	+ 9	39.92	30	42.03	49.9
37.27	+ 9	40.05	32.2	42.14	50
38.18	+12				
		(5+1)			
36.73	+ 5	37.92	15	40.50	30
37.19	9	38.11	16	41.73	35.5
37.57	12.3	39.11	25		
		(4+1)			
37.23	+ 9	37.91	41.6	40.48	31.5
39.00	16.0	37.93	41.4	40.58	18
39.07	17.7	40.45	31.1		
		(1+1)			
27.91	+95	33.93	67	36.78	50.3
31.17	84.8	34.14	67.5	37.82	41.7

Kahlenberg, 1901

%	f.t.	%	f.t.
1.90	-0.293	9.79	-1.399
2.44	-0.361	15.66	-2.591
4.87	-0.687		

Richards and Fraprie, 1901

%	f.t.
39.35	25 (5+1)
39.89	30.15 (4+1)
40.56	35

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.08	-0.194	0.82	-1.556
.16	-0.354	0.98	-1.898
.25	-0.510	1.31	-2.701
.33	-0.676	1.64	-3.668
.41	-0.792		

Caven and Johnston, 1927

39.10% f.t.= 25°

Krepelka and Rejha, 1933

%	f.t.	%	f.t.	%	f.t.
23.3	- 5	31.3	-10	32.3	11.4 E
		(7+1)			
32.4	-10	34.6	0	37.8	10
33.4	- 5	36.1	+ 5		
		(5+1)			
37.2	+ 5	38.6	20	40.4	30
37.4	10	39.5	25	41.3	35
38.2	15				
		(4+1)			
38.8	+15	39.9	30	40.8	40
39.2	20	40.4	35	41.4	45
39.5	25				
		(1+1)			
40.6	+10	36.9	45	31.3	80
40.2	15	36.3	50	30.2	85
39.6	20	35.6	55	29.0	90
39.6	25	34.9	60	27.7	95
39.3	30	34.1	65	26.2	100
38.1	35	33.2	70	26.1	100.7
37.5	40	32.2	75		

Benrath, 1941

%	f.t.	%	f.t.
40	107	20	160
35	132	15	170
30	141	10	180
25	155	5	188

Properties of phases .

Gerlach, 1867

%	d	%	d	%	d
		15°			
0	0.9991	12.87	1.128	25.74	1.284
0.68	1.006	13.55	.135	26.42	.294
1.35	.012	14.22	.143	27.09	.3027
2.03	.019	14.90	.149	27.77	.312
2.71	.024	15.58	.159	28.45	.321
3.39	.031	16.26	.165	29.13	.330
4.06	.037	16.93	.174	29.80	.339
4.74	.043	17.61	.182	30.48	.3483
5.42	.049	18.29	.189	31.16	.359
6.10	.055	18.97	.199	31.83	.369
6.73	.064	19.64	.207	32.51	.379
7.45	.071	20.32	.214	33.19	.388
8.13	.078	21.00	.223	33.87	.3974
8.81	.084	21.67	.230	34.54	.409
9.48	.092	22.35	.243	35.24	.419
10.16	.0991	23.03	.249	35.90	.429
10.84	.105	23.71	.256	36.58	.439
11.51	.113	24.38	.267	37.25	.4501
12.19	.120	25.06	.275		

Favre and Valson, 1874

m	d	m	d
23°			
0	0.998	2.0	1.262
0.5	1.071	2.5	.320
1.0	.139	3.0	.376
1.5	.202	3.5	.429

Klein, 1886

N	d	N	d
18°			
0	0.7937	3.231	1.2108
0.689	1.0456	4.257	.2756
1.476	.0982	5.321	.3400
2.034	.1343	6.639	.4187

Jahn, 1891

c	d
20°	
15.742	1.1414
27.704	1.2438

Schönrock, 1893

%	d
16°	
0	0.99900
30.8190	1.36267

Charpy, 1893

%	d	%	d
0°			
0	0.9999	11.5804	1.1238
3.0865	1.0314	14.0462	.1518
6.0172	.0621	16.7450	.1833
8.8295	.0927		

Brümmer, 1902

%	d
15°	
0	0.9990
14.410	1.1537
24.240	.2770
33.690	.4128

Herz, 1914 and 1917

N	d	N	d
25°			
0	0.9971	3.138	1.2140
0.784	1.0524	4.707	.3144
1.569	1.1073	6.276	.4099

Rohmer, 1939

%	d	%	d	
30°		34°		
38.4	1.480	38.2	1.475	
39.3	.499	39.3	.499	sic
40.0	.505	40.2	.507	
40.4	.513	41.0	.521	
38°		42°		
37.7	1.467	37.3	1.460	
39.2	.498	39.2	.497	
40.6	.510	41.0	.513	
41.7	.532	42.2	.542	

f.t.	%			
	(5+1)	(4+1)	(2+1)	(1+1)
30	40.4	40.0	39.3	38.4
34	41.0	40.2	39.3	38.2
38	41.7	40.6	39.2	37.7
42	42.2	41.0	39.2	37.3

Cabrera, Moles and Marquina, 1915

%	d	$\alpha \cdot 10^5$
20°		
0	0.9982	-
20.17	1.2288	30
27.79	1.3274	32
36.33	1.4547	26

$d^t = d^{20}\ (1-\alpha(t-20))$

Manchot, Jahrstorfer and Zepter, 1924

c	d	c	d
25°			
14.194	1.1226	29.143	1.2460
14.797	1.1283	29.594	1.2507

Flöttmann, 1928

%	t	d
37.85	15	1.4772
38.59	20	1.4866
39.554	25	1.4993

Next Page

Pesce, 1934

N	d	N	d
	25°		
1.230	1.08416	4.219	1.28083
2.406	1.11657	6.112	1.39811

Wagner, 1883

%	η (water°=100) 15°	25°	35°	45°
11.45	129.42	98.64	78.34	63.39
18.80	228.63	172.21	137.11	107.42
22.08	661.80	474.30	347.90	266.80

Herz, 1914 and 1917

N	η	N	η
	25°		
0	895	3.138	2596
0.784	1660	4.707	4854
1.569	1494	6.276	9127

Brümmer, 1902

%	σ	%	σ
	15°		
0	74.92	24.240	76.88
14.410	75.96	33.690	78.42

Jahn, 1891

c	n H_α	n D	n H_β
		20°	
0	1.3315	1.3332	1.3375
15.742	1.3558	1.3577	1.3622
27.704	1.3731	1.3751	1.3798

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.08	1.32750	0.82	1.34545
.16	.32965	0.98	.34930
.25	.33150	1.31	.35661
.33	.33362	1.64	.36357
.41	.33575		

Flöttmann, 1928

%	t	n_D
37.85	15	1.41120
38.59	20	1.41235
39.554	25	1.41453

Jahn, 1891

c	(α)magn.
	20°
0	1
15.742	1.26820
27.704	1.27663

Schönrock, 1893

%	(α)magn.
	16°
0	1
30.8190	0.3075

Jones, 1904 and Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.08	52.78	0.82	26.29
0.16	44.48	0.98	23.70
0.25	38.35	1.31	19.50
0.33	35.65	1.64	16.20
0.41	34.25		

Klein, 1886

N	κ 18°	κ 26°	$\tau . 10^4$
0.689	181.4	213.1	221
1.476	299.6	351.0	216
2.034	354.4	416.2	216
3.231	412.1	484.7	223
4.257	404.8	482.9	242
5.321	364.8	443.8	265
6.639	285.4	352.7	294

Cabrera, Moles and Marquina, 1918

%	$\chi_{mol.}$	%	$\chi_{mol.}$
	28°		
1.808	14803	10.074	14828
3.623	14752	20.17	14972
3.211	14834	27.79	15024
4.800	14796	36.33	15038
6.717	14819		

Previous Page

Water + Manganese ammonium sulfate ($MnN_2H_8O_8S_2$)

Rüdorff, 1872

%	f.t.	%	f.t.
5.66	-1.10	15.25	-3.25
9.09	-1.85	16.67	-3.55
12.28	-2.50	18.03	-4.00
13.79	-2.90		

Water + Manganese acetate ($MnC_4H_6O_4$)

Rakshit, 1925

%	d
29°	
1	1.00163
5	.01622
10	.03393
30	.10080
50	.16784

Water + Manganese benzene sulfonate ($MnC_{12}H_{10}O_6S_2$)

Ephraim and Seger, 1925

c	f.t.
13.998	18
18.840	34
24.696	49.5
32.215	64.5
42.924	80.5

Water + Ferrous chloride ($FeCl_2$)

Tammann, 1885

%	p	%	p
100°			
3.88	749.7	21.26	670.4
6.90	740.3	30.01	601.6
12.01	720.8	34.17	559.9
14.57	708.7	42.58	468.8
19.38	680.1	47.52	418.2

Lescoeur, 1894

t	p dissoc. (1+1)	p dissoc. (2+1)	p sat.sol.
40	-	-	30
50	-	-	56
60	-	-	88
70	-	-	127
80	-	-	177
90	-	30	-
100	-	48	-
110	-	75	-
120	-	130	-
125	-	188	-
155	55	-	-

Shäfer, 1949 (fig.)

t	p dissoc.	t	p dissoc.
(4+1)-(2+1)			
20	4	50	32
30	9.5	60	62.5
40	18	65	86
(2+1)-(1+1)			
50	0	90	35
60	3	100	59
70	10	110	100
80	20		
(1+1)-anh.			
120	15	150	58
130	24	160	90
140	38	170	128

Stokes, 1948

m	osmotic coefficient	m	osmotic coefficient
25°			
0.1	0.854	0.9	1.024
0.2	.863	1.0	.055
0.3	.877	1.2	.117
0.4	.896	1.4	.180
0.5	.920	1:6	.244
0.6	.943	1.8	.307
0.7	.968	2.0	.371
0.8	.994		

Etard, 1894

%	f.t.	%	f.t.
40.5	+16	45.9	53
40.9	18	49.2	72
41.0	25	51.3	89
42.5	28	51.0	96
44.4	43	51.7	118
45.0	50		

Schimmel, 1928 and Agde and Schimmel, 1928

%	f.t.		%	f.t.	
14.5	- 9		35.3	+ 7	
17.0	-11.6		36.7	8	
17.7	-13.3		37.6	12.3	tr.t.
23.3	-22.5		38.0	16.0	(4+1)
24.5	-24.0		38.6	20.5	
26.6	-29		39.2	25.0	
29.8	-35		39.4	28.5	
30.4	-36.5	E	40.4	36.5	
30.85	-38		40.9	42.5	
31.2	-39		42.6	52	
31.8	-40		43.4	56	
33.1	-43.5		43.9	60	
34.5	-47.5		44.4	63	
35.5	-50		45.5	69.5	
30.4	-36.5		45.8	70.5	
31.0	-20	(6+1)	46.6	73	
31.7	-11		47.4	76.5	tr.t.
32.1	- 6.8		47.5	82	(2+1)
32.6	- 3		47.7	86	
33.6	+ 1.5		47.9	90.5	
34.5	+ 4		48.3	96	
34.65	+ 5		50.4	117.5	

Dunn, 1902

%	d	%	d	
	15.5°			
4.96	1.0459	27.03	1.2916	
9.43	.0892	27.97	.3010	
13.02	.1255	29.95	.3289	
14.13	.1385	32.66	.3674	
16.88	.1688	35.13	.4065	
20.62	.2103	36.87	.4319	
24.07	.2528	37.53	.4439	satd.
25.00	.2625			

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.02779	4	1.2122
1	.0551	5	.2629
2	.1085	6	.3130
3	.1610		

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	69.5	4	30.84
1	60.6	5	24.05
2	48.11	6	18.13
3	38.81		

Water + Ferrous bromide ($FeBr_2$)

Lescoeur, 1894

t	p sat.sol.	t	p dissoc.
40	24	(1+1)	
50	33	160	195
60	53	(2+1)	
70	85	110	90
80	118	120	120
90	180	130	210

Etard, 1894

%	f.t.	%	f.t.
47.0	-21	56.0	37
48.3	-7	58.0	50
52.3	+10	59.4	65
53.7	21	63.3	95

Schimmel, 1929

%	f.t.		%	f.t.	
0	0		49.6	-7.0	
18.5	-6.1		50.3	-7.8	
22.8	-10.7		52.2	+12.0	
35.2	-25.0		52.3	12.5	
40.0	-36.5		53.9	21.0	
40.5	-38.5		55.4	30.0	
41.0	-40.0		57.2	41.8	
42.7	-45.0		57.6	43.0	
43.4	-47.0		58.0	46.5	
46.5	-60.0		58.4	48.5	
42.25	-43.6	E	58.45	49.0	tr.t.
43.1	-39.0	(9+1)	58.5	49.5	(4+1)
43.5	-37.0		58.6	52.0	
44.1	-34.6		58.8	57.0	
46.25	-30.7		59.5	65.0	
47.0	-30.0		61.5	75.0	
47.65	-29.3	tr.t.	63.3	83.0	tr.t.
47.7	-28.0	(6+1)	63.6	88.0	(2+1)
48.0	-22.0		64.8	100.0	
48.7	-12.5		66.6	116.0	
49.4	-9.0		68.5	124.0	
49.5	-8.0		70.2	132.0	

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.04692	72.3
1	.0932	64.3
2	.1850	50.0
3	.2756	41.94
4	.3660	33.65
5	.4560	26.25

Water (H_2O) + Sodium ferrocyanide ($Na_4FeC_6N_6$)

Conroy, 1898

%	f.t.	%	f.t.
	(10+1)		
14.31	18	29.83	60
15.18	20	35.40	77
23.20	42	37.19	80
27.06	53	38.31	96
29.12	58	38.65	98.5

Farrow, 1926

%	f.t.	d
17.04	25	1.130
26.20	50	.200
38.00	90	.290

Farrow, 1927

%	f.t.	d
17.11	24.9	1.1312
20.58	34.9	.1584
26.20	49.8	.2004
31.43	64.7	.2426
36.85	79.6	.2786
38.15	84.7	.2916
38.08	84.6	.2959
38.25	94.7	.2870
37.55	99.7	.2861

Friend, Townby and Vallance, 1929

%	f.t.	d
10.23	0.65	1.0680
14.69	16.7	1.1079
17.63	25.35	1.1303
21.54	35.75	1.1572
-	39.2	1.1665
-	42.0	1.1728
-	44.3	1.1794
24.80	46.6	-
-	49.55	1.1921
28.11	53.0	-
-	55.4	1.2078
-	59.6	1.2180
30.35	59.75	-
-	68.7	1.2446
34.34	71.0	-
35.82	74.25	-
38.26	79.5	1.2809
-	82.35	1.2892
40.69	82.4	-
41.29	84.0	-
39.57	84.2	1.2875
39.30	85.0	-
-	86.7	1.2824
39.57	88.4	-
-	94.3	1.2738
39.69	96.6	-
-	97.8	1.2688

Water + Potassium ferrocyanide ($K_4FeC_6H_6$)

Tammann, 1885

%	p
100°	
9.47	751.1
15.13	744.1
21.95	733.3
31.64	712.8
35.94	700.4

Lescoeur, 1896

t	p
(3+1)	
20	7.1
60	110.0
70	170.0
80	280.0

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.595	0.6	0.498
0.2	.556	0.7	.494
0.3	.535	0.8	.494
0.4	.518	0.9	.501
0.5	.506		

m	activity coefficient	m	activity coefficient
	25°		
0.1	(0.139)	0.6	0.056
0.2	0.100	0.7	0.052
0.3	0.081	0.8	0.048
0.4	0.070	0.9	0.046
0.5	0.062		

Bovalini and Fabris, 1937

t	p		
	(3+1)-sat.sol.	(3+1) -anh.	anh.-sat.sol.
25	19.5	10.5	-
50	84.7	54.8	-
70	218.8	175.8	-
75	267.8	233.9	-
80	330.0	298.5	-
85	415.9	394.2	-
86	434.2	419.8	-
87.3	458.0	458.0	-
88	-	-	473.0
90	-	-	510.0
95	-	-	596
100	-	-	679
104.2(b.t.)	-	-	760

Etard, 1894

%	f.t.	%	f.t.
10.8	-2	34.0	60
15.4	+7	39.1	75
17.9	14	41.9	89
23.0	30	42.6	98
31.7	56	46.8	157

Jones, 1904; Jones and Bassett, 1905

M	f.t.
0.1	-0.580
.2	-1.05
.3	-1.45
.4	-1.78

Fabris, 1921

%	f.t.	%	f.t.
0	0	12.48	0
1.23	-0.24	17.95	10
1.96	-0.36	23.36	23
3.85	-0.62	27.79	35
5.66	-0.86	29.28	40
7.41	-1.09	32.64	50
9.09	-1.30	35.55	60
10.71	-1.50	38.07	70
E	-1.58	39.20	75
		40.29	80
		41.65	85
		43.17	90
		44.49	94
		45.82	99
		46.07	100

Vallance, 1922

%	f.t.	%	f.t.
17.541	10.4	21.953	20.4
19.067	13.8	23.074	23.1
20.862	16.9	23.971	25.0

Farrow, 1927

%	f.t.	%	f.t.
23.95	24.9	36.94	64.7
28.01	34.9	40.45	79.6
33.13	49.8	43.78	99.7

Fabris, 1932

%	f.t.	%	f.t.	%	f.t.
		I		II	
38.19	70	41.38	90	41.68	88.3
39.02	75	41.62	92	42.24	90
40.07	80	41.99	94	43.58	92.6
40.22	82	42.19	96	43.43	92.8
40.33	83	42.58	99.4	43.91	94
40.70	84	42.63	99.6	44.74	95.8
40.72	85	43.13	100		
40.94	86	44.77	104.1		
40.98	86.5				
41.34	87.3				
	tr.t.				

Bovalini and Fabris, 1933

%	f.t.	%	f.t.
	758 mm		
12.48	0(3+1)	38.75	74
25.94	30	40.72	85
29.28	40	41.38	90 B
36.81	65	44.77	104.1 b.t.

Schiff, 1860

%	d	%	d
	15°		
0	0.9994	11	1.0659
1	1.0049	12	.0724
2	.0107	13	.0790
3	.0166	14	.0856
4	.0225	15	.0922
5	.0286	16	.0989
6	.0347	17	.1057
7	.0408	18	.1026
8	.0470	19	.1195
9	.0532	20	.1205
10	.0595		

Jones, 1904; Jones and Bassett, 1905

M	d
0°	
0.1	1.027176
.2	.051100
.3	.074976
.4	.110084
.5	.122160

Chéneveau, 1907

%	d
18.2°	
0	0.9982
13.23	1.0916

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.0307	1.5	1.0896
1	1.0605	2	1.1182

Rakshit, 1925

%	d	%	d
	20°		
1	1.00477	10	1.05658
5	1.02818	30	1.16472

Farrow, 1927

%	t	d
	sat.sol.	
23.95	24.9	1.1731
28.01	34.9	.2018
33.13	49.8	,2350
36.94	64.7	,2635
40.45	79.6	.2854
43.78	99.7	.3115

Chéneveau, 1907

%	n C	n D	n Tl	n F	n G'
			18.2°		
0	-	1.33313	-	-	-
13.23	1.35879	.36099	1.36332	1.36608	1.37031

Jones, 1904 and Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.1	171.8	0.4	156.0
0.2	163.4	0.5	154.1
0.3	158.4		

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	78.3	1.5	64.5
1	66.8	2	62

Dewar and Fleming, 1897

t	ε	t	ε
		12%	
-204.3	35.1	-129.5	68.5
-203.8	34.9	-119.4	81.7
-200.0	36.4	-111.4	97.0
-187.2	39.8	-107.0	115.0
-168.1	42.4	-105.3	126.0
-156.5	44.3	-102.0	147.0
-145.8	49.8	-98.8	168.0

Water + Calcium ferrocyanide ($Ca_2FeC_6N_6$)

Berkeley, Hartley and Stephenson, 1909

%	P osmotic	d
	0°	
0	-	0.99987
6.60	5.34	1.05716
10.88	9.20	.19592
14.97	14.65	.13444
17.90	20.33	.16287

Farrow, 1926 and 1927

%	f.t.	d^t
30.45	-10.1	-
36.44	+24.9	1.3563
39.22	34.9	.3662
42.04	49.8	.3970
44.20	59.7	-
44.44	64.7	1.4074
44.37	79.6 and 90	-

Water + Strontium ferrocyanide ($Sr_2FeC_6N_6$)

Berkeley, Hartley and Stephenson, 1909

%	P	d
	0°	
0	-	0.99987
5.82	3.40	1.05291
11.51	6.18	.10892
15.23	8.59	.14690
19.34	12.04	.19253

Water + α-Tetramethyl ferrocyanide ($C_{10}H_{12}N_6Fe$)

Berkeley and Hartley, 1915 - 1916

%	osmotic P	d
	0°	
0	-	0.99987
6.80	5.96	1.01843
36.92	52.32	1.10632

Water + Ferrous iodide (FeI_2)

Lescoeur, 1894

t	p dissoc. (1+1)	p dissoc. (2+1)	p sat.sol.
50	-	-	30
60	-	-	50
70	-	-	74
80	-	-	101
90	-	77	144
100	-	115	-
110	-	160	-
120	-	233	-
160	110	-	-

Water + Ferrous nitrate (FeN_2O_6)

Funk, 1899

%	f.t.	%	f.t.
-14.5	29.76	-9 (6+1)	39.68
-19	32.36	0	41.53
-21	33.33	+18	45.14
-27 (9+1)	35.66	24	46.51
-21.5	36.10	60.5	62.50
-19	36.56		
-15.5	37.17		

Traube, 1895

%	d
15°	
0	0.99913
10.257	1.06110
15.598	.13466
23.962	.22022

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.03111	76.0
1	.0617	63.8
2	.1223	48.53
3	.1815	37.50
4	.2412	28.50

Water + Ferrous perchlorate ($FeCl_2O_8$)

Lindstrand, 1936

%	f.t.	d
63.39	0	1.543
66.93	20	.560
67.76	25	.565
68.67	30	.569
69.48	35	.573
70.34	40	.577
71.05	42	.569
71.55	45	.571
72.19	50	.574
72.84	55	.577
73.49	60	.580

Water + Ferrous sulfate ($FeSO_4$)

Tammann, 1885

t	p 0%	p 18.60%	p 30.26%
46.13	76.2	73.7	70.7
53.90	112.0	108.7	104.2
55.83	122.9	119.1	114.1
59.29	144.6	140.5	134.6
63.65	176.6	171.5	165.3
63.70	177.0	172.0	165.5
67.22	207.1	201.1	193.6
70.06	234.4	226.8	219.4
73.85	275.6	267.8	259.3
77.70	323.6	314.5	304.4
80.39	361.0	349.6	339.0
83.49	408.8	397.1	385.2
86.69	463.2	450.6	437.4
89.67	519.5	505.0	489.7
82.28	573.1	559.0	541.3
94.74	627.9	610.2	597.9
97.02	682.6	663.6	645.9
99.99	759.5	740.1	719.3

%	p	%	p
	100°		
8.13	753.4	22.10	739.7
13.76	749.0	31.95	716.3
19.34	743.5	34.75	710.2

Nicol, 1884

t	p	t	p
sat.sol.		satd at 95°	
65	169.9	65	165.7
75	263.5	75	259.3
85	394.7	85	392.7
95	587.8	95	587.8

Pearce and Eckstrom, 1937

M	p	M	p
	25°		
0.0	23.752	0.4	23.399
.1	23.656	.6	23.239
.2	23.565	.6889	23.172

Wiedemann, 1874

t	p dissoc.	t	p dissoc.
25	15.6	55	90.1
30	19.9	60	114.1
35	28.1	65	140.8
40	37.1	70	175.9
45	50.3	75	222.1
50	58.1	80	267.9

Cohen, 1900

t	p dissoc.	t	p dissoc.
30.67	21.76 (7+1)	44.45	52.86
39.96	39.94	46.43	59.63

Gerlach, 1886

%	b.t.	%	b.t.
0	100	33.51	101.5
15.03	100.5	34.72	101.6
25.59	101		

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
3.14	100.10	19.64	100.65
6.28	100.19	21.04	100.73
8.44	100.25	22.35	100.83
11.57	100.35	23.49	100.92
13.65	100.42	24.68	101.02
15.22	100.49	26.12	101.12
18.57	100.56		

Tobler, 1855

%	f.t.	%	f.t.
13.6	0	24.6	30
16.6	10	26.7	37
17.4	12	30.0	45
20.6	20	32.0	55
21.5	21	36.1	70

Mulder, 1864

%	f.t.	%	f.t.
7.32	0	38.98	75.5
17.89	16.5	38.00	83.5
26.90	33	31.60	98
30.89	42.5	30.79	99
39.21	63 tr.t.		

Guthrie, 1876

%	f.t.	%	f.t.
5	-0.2	14.5	-2.0 E
10	-0.8	14.9	0.0 (7+1)

Etard, 1894

%	f.t.	%	f.t.
13.0	-1	37.8	77
15.1	+5	38.7	86
22.7	24	36.7	94
26.3	34	34.7	102
32.5	52	28.0	112
36.4	60	17.3	130
37.7	67	2.5	152

Koppel, 1905

E: -2°

Fraenkel, 1907

%	f.t.	%	f.t.	%	f.t.
1.0	-0.172	8.01	-1.063	12.27	-1.717
2.55	-0.370	8.81	-1.153	12.66	-1.771
2.56	-0.373	9.46	-1.247	12.92	-1.819
2.76	-0.419	10.18	-1.360	14.18	-1.820
4.11	-0.566	11.22	-1.465	15.03	-1.824
4.91	-0.651	11.27	-1.511	15.45	-1.812
5.17	-0.685	11.88	-1.615	18.30	-1.821
6.39	-0.820	12.00	-1.637	18.99	-1.828
7.22	-0.980				
		(7+1)			
13.54	0	22.84	25.02	30.71	45.18
17.06	+10.00	24.77	30.03	32.71	50.21
19.26	15.25	26.84	35.07	33.42	52.00
20.98	20.13	28.67	40.05	34.24	54.03
		(4+1)			
35.46	60.01	34.35	68.02	28.80	85.02
35.73	65.00	31.46	77.00	27.15	90.13
35.93	70.04	30.36	80.41		

E : 14.98 % -1.821° tr.t. (7+1)-(4+1) 56.7°

(7+1)-(1+1) 64.0°

Agde and Barkholt, 1926

%	f.t.	%	f.t.
13.81	1	26.61	34
17.14	9.6	30.05	43
21.34	21	34.32	54
23.02	25	30.53	80

Properties of phases .

Schiff, 1859 - 1860

%	d	%	d
	17.2°		
0.55	1.0040	11.49	1.1200
1.09	.0092	12.03	.1264
1.64	.0145	12.58	.1329
2.19	.0199	13.12	.1394
2.73	.0253	13.67	.1459
3.28	.0308	14.22	.1525
3.83	.0364	14.77	.1592
4.37	.0420	15.31	.1658
4.92	.0477	15.86	.1793
5.47	.0534	16.40	.1861
6.02	.0592	16.95	.1930
6.56	.0651	17.49	.1999
7.11	.0710	18.04	.2039
7.66	.0769	18.59	.2139
8.20	.0829	19.14	.2210
8.75	.0890	19.69	.2281
9.30	.0950	20.24	.2353
9.84	.1012	20.78	.2425
10.38	.1074	21.33	.2498
10.94	.1137		

%	d
17.2°	
0	0.9988
2.278	1.0207
4.556	1.0437
6.835	1.0680
9.112	1.0929
13.670	1.1459
20.505	1.2317

Gerlach, 1867

%	d	%	d
	15°		
0	0.999		
0.55	1.004	12.03	1.124
1.10	.010	12.58	.130
1.64	.015	13.12	.136
2.18	.020	13.67	.142
2.73	.026	14.22	.148
3.28	.031	14.76	.154
3.83	.036	15.30	.160
4.38	.042	15.85	.167
4.92	.047	16.40	.173
5.47	.053	16.95	.179
6.02	.058	17.49	.186
6.57	.064	18.04	.192
7.11	.070	18.58	.199
7.65	.076	19.13	.205
8.20	.081	19.68	.212
8.74	.087	20.23	.218
9.29	.093	20.77	.225
9.84	.099	21.32	.231
10.39	.105	21.87	.238
10.94	.111		
11.48	.117		

Klein, 1886

N	d	N	d
	18°		
0	0.7937	2.0	1.1375
0.5	1.0344	3.0	.2018
1.0	1.0692	3.5	.2359

Brümmer, 1902

%	d
15°	
0	0.9993
3.647	1.0358
7.181	.0716
11.661	.1205
17.906	.1917

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
10.938	1.1017
21.845	1.2014

Rakshit, 1925

c (7+1)	d
20°	
1	1.00363
5	.02344
10	.04819
30	.14478
50	.23865

Agde and Barkholt, 1926

%	d	%	d
	sat.sol.		
13.81	1.140	26.61	1.312
17.14	.178	30.05	.363
21.34	.233	30.53	.367
23.02	.255	34.32	.432

Pearce and Eckstrom, 1937

M	d	M	d
	25°		
0.0	0.997074	0.4	1.049670
.1	1.010772	.6	.074320
.2	1.023981	.6889	.085004

Holtzman and May, 1951

c	d
18°	
9.87	1.097
10.84	.105
11.78	.113
12.61	.122
13.35	.131

Schmidt, 1859

t	π	d
16.5	38.0	1.104
16.3	41.3	1.078
15.4	43.1	1.050
15.6	44.6	1.027

Kasankin, 1891

d	capillary rise	d	capillary rise
16-17°		32.34°	
1.225	22.41	1.3045	21.33

Brümmer, 1902

%	σ
15°	
0	74.92
3.647	68.67
7.181	67.36
11.661	69.34
17.906	76.55

Forch, 1905

N	t	σ	N	t	σ
0.600	16.7	77.15	1.515	17.2	77.86
1.010	17.0	77.44	2.000	16.7	78.29
1.241	16.9	77.60	2.439	17.1	78.62
			3.000	16.6	79.28

Klein, 1886

N	κ 18°	κ 26°
0.5	146.9	172.6
1.0	245.4	288.1
2.0	371.2	437.4
3.0	371.2	519.8
3.5	447.8	534.7

Gray, 1879-1880

d	U
at room temp.	
1.146	0.8814
1.1523	0.8468

Agde and Holtmann, 1926

%	U	%	U
25-45°			
0.55	0.999	18.5	0.806
1.10	.973	19.1	.802
1.65	.963	19.6	.799
2.20	.953	20.1	.795
2.73	.945	20.7	.791
3.3	.933	21.2	.787
3.9	.930	21.8	.783
4.4	.923	22.3	.780
5.0	.915	22.9	.777
5.5	.910	23.4	.773
6.0	.903	23.9	.770
6.6	.898	24.5	.767
7.1	.893	25.0	.765
7.7	.888	25.6	.760
8.2	.883	26.1	.757
8.7	.878	26.7	.752
9.3	.875	27.3	.750
9.8	.870	27.8	.745
10.4	.865	28.3	.742
10.9	.860	28.8	.740
11.4	.856	29.4	.735
11.9	.853	30.0	.732
12.6	.848	30.5	.730
13.1	.845	31.1	.727
13.7	.840	31.6	.722
14.2	.835	32.0	.720
14.7	.833	32.8	.717
15.3	.828	33.3	.715
15.8	.825	33.9	.712
16.4	.820	34.4	.707
16.9	.817	35.0	.705
17.4	.813	35.5	.702
18.0	.810		

Kobe and Cooch, Jr., 1954

t	enthalpie, cal./gr. 0%	5%	10%	15%	20%	25%	30%	35%
0	0.00	-4.2	-8.8	-	-	-	-	-
5	5.03	0.4	-4.5	-9.4	-	-	-	-
10	10.04	5.0	-0.1	-5.2	-	-	-	-
15	15.03	9.7	4.3	-1.1	-	-	-	-
20	20.02	14.4	8.7	3.1	-2.6	-	-	-
25	25.01	19.0	13.1	7.2	1.4	-	-	-
30	30.00	23.7	17.5	11.4	5.4	-	-	-
35	34.99	28.4	22.0	15.6	9.4	3.2	-	-
40	39.98	33.0	26.4	19.8	13.4	7.0	-	-
45	44.97	37.7	30.8	24.0	17.4	10.8	4.3	-
50	49.96	42.4	35.3	28.3	21.4	14.7	8.0	-
55	54.94	47.1	39.7	32.5	25.4	18.5	11.7	-
60	59.93	51.8	44.2	36.8	29.5	22.4	15.3	8.4
65	64.94	56.5	48.7	41.0	33.5	26.2	19.0	11.9
70	69.94	61.2	53.2	45.3	37.6	30.1	22.7	-
75	74.94	66.0	57.7	49.6	41.7	34.0	26.4	-
80	79.95	70.7	62.2	53.9	45.8	37.9	30.2	-
85	84.96	75.4	66.7	58.2	49.8	41.5	-	-
90	89.98	80.2	71.2	62.5	54.0	45.7	-	-
95	95.01	84.9	75.8	66.8	58.1	49.7	-	-
100	100.04	89.7	80.3	71.2	62.2	-	-	-

Water + Ammonium ferrous sulfate ($FeN_2H_8O_8S_2$)

Tobler, 1855

%	f.t.	%	f.t.
10.9	0	26.6	45
14.9	12	28.7	55
17.8	20	30.8	60
21.9	30	33.2	65
24.1	36	36.2	75

Schiff, 1859-1860

%	d	%	d
	19°		
0	0.998	11.60	1.097
0.72	1.004	12.32	.104
1.45	.011	13.05	.110
2.17	.016	13.77	.116
2.90	.022	14.49	.124
3.62	.028	15.21	.130
4.35	.034	15.94	.136
5.07	.040	16.66	.143
5.80	.045	17.39	.150
6.52	.052	18.11	.156
7.25	.058	18.84	.164
7.97	.064	19.56	.171
8.70	.071	20.29	.179
9.42	.078	21.02	.185
10.15	.083	21.74	.193
10.87	.090		

Rakshit, 1925

c (6+1)	d
20°	
1	1.00392
5	.02673
10	.05454
30	.15968

Water + Sodium ferrous sulfate ($FeNa_2O_8S_2$)

Koppel, 1905

%		f.t.
31.34		24.92
31.34		35
31.79		40
31.89	(4+1)	21.8

Water + Potassium ferrous sulfate ($K_2FeS_2O_8$)

Tobler, 1855

%	f.t.	%	f.t.	
16.4	0	29.1	35	
20.3	10	31.0	40	
22.5	14.5	35.9	55	(6+1)
23.6	16	37.2	65	
26.7	25	39.1	70	

Küster and Thiel, 1899

Solubility of hydrates at 0-95°

tr.t. (6+1)-(4+1) 30° (4+1)-(2+1) 87°

(6+1)-(2+1) 54°

Water + Nickel chloride ($NiCl_2$)

Tammann, 1885

%	p	%	p
	100°		
6.10	743.9	25.35	638.8
10.55	728.1	26.50	624.0
17.06	697.3	31.89	574.7
21.93	664.7	32.77	561.3
22.02	665.0	38.92	489.9

Dieterici, 1898

m	p
0°	
0.000	4.579
0.994	4.324
2.006	3.922
2.572	3.538

Derby and Yngve, 1916

t	p (6+1)	p (4+1)	p dissoc. (6+1)-(4+1)	p dissoc. (4+1)-(2+1)
19.8	9.6	-	7.3	-
20.6	-	-	7.8	-
24.1	12.0	-	10.1	-
24.6	-	-	10.4	-
25.95	-	-	-	6.0
26.0	-	-	-	6.1
30.3	-	-	15.2	-
30.7	-	-	15.2	-
31.0	17.5	-	16.1	8.9
32.31	-	-	-	-
35.05	21.5	-	20.9	-
36.25	22.5	22.5	22.5	-
38.20	-	-	-	14.2
39.80	26.6	-	-	-
40.57	-	28.9	-	-
45.22	33.3	36.7	-	-
47.69	-	-	-	26.4
48.34	-	42.6	-	-
53.1	-	56.4	-	-
54.5	-	-	-	40.4
54.63	-	-	-	40.7
59.63	-	-	-	56.3
66.34	-	-	-	84.1
74.06	-	-	-	108.1

Lescoeur, 1890

t	p dissoc. (6+1)	p dissoc. (2+1)
15	3.4	-
20	4.6	-
25	6.3	-
30	10.5	-
40	24.0	-
100	-	125

Isopiestic solutions

Robinson and Stokes, 1940

m_1	m_2	m_1	m_2
	25°		
0.1188	0.1662	0.1975	0.2820
0.7864	1.328	0.9430	1.661
1.831	3.942	2.123	4.81
0.2158	0.3113	0.5380	0.8482
1.212	2.284	1.443	2.872

1 - nickel chloride

2 - potassium chloride

Stokes, 1948

m	osmotic coefficient
25°	
0.1	0.857
.2	.868
.3	.885
.4	.907
.5	.934
.6	.960
.7	.987
.8	1.016
.9	.048
1.0	.082
.2	.150
.4	.221
.6	.293
.8	.366
2.0	.442
2.5	.633
3.0	.816
3.5	.969
4.0	2.100
4.5	.202
5.0	.292
5.8	.407

Rüdorff, 1872

%	f.t.	%	f.t.
4.41	-2.00	15.47	-8.85
9.02	-4.40	18.02	-13.90
12.48	-7.25	20.29	-17.10

Etard, 1894

%	f.t.	%	f.t.
29.7	-17	41.9	+38
31.0	-16	45.0	59
37.3	+10	46.6	78
38.5	+18	46.7	96

Jones, 1904; Jones and Bassett, 1905; Jones and Getman, 1904

M	f.t.	M	f.t.
0.037	-0.205	0.800	-4.880
.074	-0.380	0.900	-5.740
.149	-0.768	1.000	-6.580
.223	-1.170	1.500	-11.920
.297	-1.585	2.000	-20.000
.372	-2.032	2.500	-31.500
.446	-2.458	3.000	-41.500
.521	-2.945	3.483	-53.000
.743	-4.547		

Boye, 1933

%	f.t.		%	f.t.	
0	0		34.0	-14.9	
2.17	-0.5		34.1	-9.8	
3.54	-1.5		34.3	-5.4	
7.18	-3.6		34.8	0	
12.0	-7.7		40.4	+26.3	
15.6	-11.3		41.3	27.9	
20.6	-19.0		41.5	28.5	
22.0	-22.1		41.6	28.8	tr.t.
24.2	-26.7		41.6	29.5	(4+1)
26.9	-35.3		41.7	31.2	
29.0	-41.0		42.0	35.0	
29.6	-44.5		45.3	61.6	
29.9	-45.3	E	45.9	63.8	
29.9	-44.0	(7+1)	46.0	64.0	
30.0	-40.8		46.1	64.3	tr.t.
30.1	-38.0		46.1	65.8	(2+1)
30.3	-36.1		46.1	67.6	
31.3	-34.2		46.3	73.2	
32.7	-33.5		46.7	100.2	
33.8	-33.3	tr.t.	46.8	110.4	
33.8	-32.4	(6+1)	46.8	112.2	
33.8	-29.1		46.9	117.9	
33.9	-20.2				

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987		
1	1.0086	13	1.1330
2	.0185	14	.1446
3	.0284	15	.1563
4	.0383	16	.1695
5	.0480	17	.1838
6	.0581	18	.1961
7	.0681	19	.2094
8	.0781	20	.2229
9	.0881	21	.2380
10	.0981	22	.2530
11	.1099	23	.2681
12	.1215	24	.2832
		25	.2985

Favre and Valson, 1874

m	d	m	d
	23.1°		
0	0.998	2.0	1.230
0.5	1.061	2.5	.284
1.0	.119	3.0	.335
1.5	.176		

Cabrera, 1915

%	d	$a.10^6$	$b.10^7$
	15°		
2.21	1.02024	174	45
4.87	.04631	198	28
15.84	.16365	281	15
23.79	.26142	302	8
34.63	.42188	334	4

$d^t = d^{15}\ (1-a(t-15)-b(t-15)^2)$

Heydweiller, 1921

N	d
18°	
0.5	1.03054
1	.0604
2	.1188
3	.1757
4	.2316

Wagner, 1883

%	η (water°=100) 15°	25°	35°	45°
11.449	90.40	70.03	57.46	48.25
22.690	140.20	109.70	87.84	72.71
30.400	229.50	171.80	139.20	229.50

Jones, 1904; Jones and Bassett, 1905; Jones and Getman, 1904

M	d	M	d
	0°		
0.037	1.003340	0.800	1.094304
.074	.007440	0.900	.103812
.149	.016792	1.000	.117344
.223	.025720	1.500	.174924
.297	.033800	2.000	.231744
.372	.043680	2.500	.282856
.446	.051816	3.000	.341552
.521	.059144	3.483	.391772
.743	.087352		

Jones, 1904; Jones and Bassett, 1905; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.037	1.3272	0.372	1.33684
.074	.32790	.446	.33894
.149	.32965	.521	.34106
.223	.33240	.743	.34710
.297	.33485		

Heydweiller and Grube, 1916

w.l. (in cm.10^5)	Dn . 10^5 (sol. - aq.) 0.4960 N	1.002 N	2.037 N	4.014 N
		18°		
2.314	1055	2094	4140	-
2.574	956	1899	3756	7167
2.749	915	1817	3596	6849
2.981	878	1739	3445	6553
3.256	847	1679	3325	6329
3.405	835	1655	3276	6332
4.679	783	1548	3064	5799

Jones, 1904; Jones and Bassett, 1905; Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.037	103.20	0.800	61.87
.074	97.00	0.900	59.19
.149	88.64	1.000	58.23
.223	85.50	1.500	46.85
.297	82.52	2.000	38.59
.372	77.90	2.500	29.88
.446	77.20	3.000	22.85
.521	75.20	3.483	16.92
.743	70.23		

Heydweiller, 1921

N	λ
18°	
0.5	70.8
1	62.1
2	50.6
3	41.0
4	33.33

Marignac, 1876

%	U
24-55°	
3.48	0.9451
6.72	.9017
12.61	.8310
27.05	.7351

Kapustinskii, Yakuchevskii and Drakin, 1953

%	U	%	U
0	0.9980	13.38	0.8220
4.36	.9371	17.36	0.7752
6.54	.9086	20.90	0.7370
8.89	.8785		

Water + Nickel bromide ($NiBr_2$)

Etard, 1894

%	f.t.	%	f.t.
47.1	-21	60.3	+77
51.7	-6	61.0	98
56.6	+19	61.0	100
58.9	38	60.7	140
60.5	58		

Heydweiller, 1921

N	d	λ
	18°	
0.1	1.00989	86.3
.2	.01978	81.1
.5	.04922	73.4
1	.0981	65.8
2	.1945	54.54
3	.2897	45.1
4	.3840	37.48

Water + Nickel iodide (NiI_2)

Etard, 1894

%	f.t.	%	f.t.
51.8	-23	64.1	43
54.3	-6	65.0	80
57.8	+11	65.2	85
59.0	16	65.7	90

Water + Nickel chlorate ($NiCl_2O_6$)

Meusser, 1902

%	f.t.	%	f.t.
-9	26.62	+48.5 (4+1)	67.60
-13.5	31.85	55	68.78
-18 (6+1)	49.55	65	69.05
-8	51.52	79.5	75.50
0	52.66		
+18	56.74		
40	64.47		

d sat.sol. 18° 1.661

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.04512	62.0
1	.0895	54.3
2	.1772	43.67
3	.2630	34.7
4	.3470	27.50

Water + Nickel nitrate (NiN_2O_6)

Tammann, 1885

%	p	%	p
	100°		
9.21	741.8	28.65	657.7
12.56	731.4	29.70	649.0
14.23	726.4	38.11	578.1
19.88	705.9	41.13	538.5
23.26	687.7	44.52	504.2

Dieterici, 1898

m	p
0°	
0	4.579
1.001	4.363
1.969	4.045
2.981	3.599
3.469	3.202

Rüdorff, 1872

%	f.t.
5.69	-1.60
10.47	-3.40
14.50	-5.35
17.95	-7.40

Funk, 1899

%	f.t.	%	f.t.
30.63	-15	44.32	0
33.58	-18.5	49.06	+20
37.29	-23	55.22	41
39.02	-23 (9+1)	62.76	56.7
39.48	-21	61.61	58 (3+1)
44.13	-10.5	61.99	60
39.94	-21 (6+1)	62.76	64
41.59	-12.5	63.95	70
42.11	-10	70.16	90
43.00	-6	77.12	95

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	f.t.	M	f.t.
0.0761	-0.377	1.0654	-6.691
.1522	-0.750	1.2176	-8.006
.3044	-1.527	1.5220	-10.883
.6088	-3.273	2.0000	-17.000
.7610	-4.306	2.5000	-25.000

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.01	-0.0550	0.25	-1.251
.025	-0.1299	.5	-2.652
.05	-0.2487	.75	-4.213
.075	-0.3664	1.0	-6.101
.10	-0.4960	1.5	-10.576
		2.0	-17.050

Sieverts and Schreiner, 1934

%	f.t.	%	f.t.
8.7	- 1.6	42.6	-13.1
16.2	- 6.0	43.6	- 2.9
21.7	- 9.4	44.2	0.0
22.3	-10.3	48.5	+20.0
25.0	-12.5	50.0	25.0
27.4	-15.0	51.3	30.0
28.3	-16.0	54.3	40.0
30.9	-19.4	58.2	50.0
32.2	-21.3	60.0 tr.t.	54.0
33.4	-22.1	60.5 (4+1)	55.0
36.0 E	-27.8	60.6	55.8
36.2 (9+1)	-27.2	61.2	60.0
36.7	-25.4	64.3	75.4
37.0	-24.9	65.6	80.6
37.2	-23.7	67.2 tr.t.	85.4
38.3	-20.0	68.2 (2+1)	90.4
40.2	-13.8	68.0	94.2
41.2	-11.1	68.2	95.0
38.7 E ice+(6+1)	-34.1	69.2	99.5
40.0 (6+1)	-25.9	69.7	110.5
40.3	-24.9	71.6	116.0
40.5	-20.6	70.6	119.8
41.2	-18.1		

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987	21	1.2040
1	1.0079	22	.2160
2	.0171	23	.2280
3	.0263	24	.2400
4	.0355	25	.2517
5	.0450	26	.2651
6	.0537	27	.2781
7	.0626	28	.2911
8	.0715	29	.3041
9	.0804	30	.3177
10	.0889	31	.3314
11	.0985	32	.3455
12	.1080	33	.3595
13	.1174	34	.3736
14	.1268	35	.3878
15	.1360	36	.4031
16	.1475	37	.4185
17	.1587	38	.4338
18	.1699	39	.4490
19	.1810	40	.4648
20	.1919		

Favre and Valson, 1874

m	d	m	d
	24.4°		
0	0.997	2	1.266
0.5	1.073	2.5	.324
1	1.141	3	.378
1.5	1.205		

Traube, 1895

%	d	%	d
	20°		
0	0.99823	13.497	1.12481
4.400	1.03663	19.673	1.19124
6.948	1.06051	30.258	1.32017

Cabrera, 1914

%	d	$a.10^6$	$b.10^7$
	15°		
0	0.99913	-	-
1.805	1.01428	-182	-46
4.515	.03778	223	40
4.78	.04005	228	32
12.19	.11084	324	22
13.99	.12817	344	18
24.51	.24479	435	19
24.49	.24485	-444	-11

$d^t=d^{15}(1-1(t-15)-b(t-15)^2)$

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	d	M	d
	0°		
0	0.999868	1.0654	1.157136
0.0761	1.009256	1.2176	.177832
.1522	.022360	1.5220	.224284
.3044	.044528	2.0000	.294220
.6088	.091960	2.5000	.363728
.7610	.111896	3.0000	.428384
		3.3940	.478048

Jones and Pearce, 1907

M	d	M	d
	20°		
0	0.99823	0.25	1.03653
0.01	0.99974	.50	.07420
.025	1.00210	.75	.11113
.05	.00600	1.00	.14360
.075	.00974	1.50	.22919
.10	.01350	2.00	.29429

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.0379	4	1.2920
1	.0754	5	.3607
2	.1492	6	.4284
3	.2211		

Wagner, 1883

%	η			
	15°	25°	35°	45°
40.9530	222.60	169.70	128.20	108.30
30.0060	135.60	105.90	85.51	70.69
16.493	90.72	70.13	57.42	48.87

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	n_D	M	n_D
	0°		
0.0761	1.32782	0.7610	1.34667
.1522	.32989	1.0654	.35463
.3044	.33378	1.2176	.35870
.6088	.34277	1.5220	.36653

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.0761	90.90	1.0654	55.30
0.1522	85.14	1.2176	42.28
0.3044	76.00	1.5220	45.29
0.6088	68.60	2.0000	34.25
0.7610	61.90	2.5000	26.11

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.010	106.77	0.50	71.44
0.025	100.31	0.75	64.06
0.050	93.57	1.00	57.38
0.075	89.74	1.50	45.79
0.10	87.84	2.00	35.61
0.25	80.07		

Heydweiller, 1921

N		N	
0.5	66.9	4	29.3
1	57.6	5	22.7
2	46.0	6	16.0
3	37.2		

Marignac, 1876

%	U
24-55°	
4.80	0.9409
9.22	0.8949
16.88	0.8228
28.91	0.7171

Water + Nickel perchlorate ($NiCl_2O_8$)

Terlikowski and Golblum, 1911 and Terlikowski, 1912

c	f.t.	d
33.05	-10.9	-
46.61	-21.3	-
69.80	-30.7	-
89.98	-30.7	-
92.48	-21.3	-
104.55	0	1.5726
106.76	+ 7.5	1.5755
110.05	+18	1.5760
112.15	+26	1.5841
118.60	+45	1.5936

E : -49.0°

Water + Nickel sulfate (NiO_4S)

Heterogeneous equilibria.

Wüllner, 1860

t	p 0%	p 9.09%	p 16.67%
48.69	86.21	84.48	83.09
51.70	100.07	98.44	96.12
58.63	140.06	137.09	134.85
60.80	153.29	150.37	145.49
62.28	165.45	162.19	158.98
65.64	191.07	187.12	183.36
68.45	217.90	213.71	209.90
70.60	239.27	235.13	230.64
74.38	281.90	277.67	272.24
78.80	337.74	332.03	326.10
82.63	393.83	386.43	379.58
84.80	429.51	421.46	413.64
86.50	459.21	450.56	443.20
89.93	524.15	510.82	505.10
92.20	571.03	560.51	550.38
94.65	625.63	614.32	603.01
96.85	678.31	666.00	653.60
99.30	741.28	728.08	715.38

Tammann, 1885

t	p 0%	p 20.32%	p 30.27%
25.65	24.7	24.2	22.6
33.77	39.4	38.5	36.6
41.12	61.9	60.5	57.3
46.94	79.4	77.4	73.4
53.07	107.6	105.0	100.4
60.60	154.0	151.1	143.2
65.06	188.1	182.8	175.6
68.65	220.5	213.8	204.5
72.62	261.6	254.4	243.9
75.61	296.8	288.8	277.9
79.20	344.1	334.3	322.2
82.92	399.6	388.8	375.2
86.46	459.1	446.3	430.9
90.13	528.6	514.4	496.5
93.73	604.9	588.5	569.4
97.01	682.4	665.6	643.4
100.11	763.1	742.9	723.1

%	p
100°	
10.16	752.6
16.64	747.0
20.89	742.6
25.62	735.0
28.09	727.4

Lescoeur, 1895

t	p sat.sol.
15	10.5
20	14.4
60	160
75	270

Dieterici, 1898

m	p
	0°
0	4.579
0.467	.553
0.870	.529
1.516	.482

Wiedemann, 1874

t	p dissoc.	t	p dissoc.
25	18.3	55	92.1
30	25.3	60	115.0
35	33.0	65	143.0
40	43.6	70	175.3
45	57.2	75	211.0
50	63.1	80	252.2

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.581	0.9	0.456
.2	.533	1.0	.459
.3	.508	.2	.472
.4	.488	.4	.492
.5	.475	.6	.517
.6	.465	.8	.551
.7	.458	2.0	.589
.8	.456	2.5	.708

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
2.70	100.10	20.33	100.86
4.99	100.17	21.36	100.96
7.13	100.24	22.50	101.07
10.07	100.34	23.75	101.22
13.14	100.46	24.64	101.33
15.21	100.55	25.63	101.42
17.24	100.66	26.02	101.62
18.79	100.76	27.40	101.78

Chrétien and Rohmer, 1934

sat.sol.	b.t.	sat.sol.	b.t.
(1+1)	103.9	(4+1)	106.6
(2+1)	105.1	(5+1)	107.2
(3+1)	105.9	(6+1)	107.7

Tobler, 1855

%	f.t.	%	f.t.
23.3	2	32.9	41
27.2	16	34.2	50
28.4	20	35.2	53
29.1	23	36.4	60
31.2	31	38.2	70

Mulder, 1864

%	f.t.	%	f.t.
22.63	0	36.31	54
25.98	14	40.19	72
27.90	20.5	45.20	99.4
32.89	40.25	47.01	108.4 b.t.

Rüdorff, 1872

%	f.t.	%	f.t.
6.76	-0.80	14.59	-2.05
7.60	-0.90	15.74	-2.25
9.93	-1.20	17.35	-2.55
12.16	-1.60		

Etard, 1894

%	f.t.	%	f.t.
21.7	-3	38.7	74
22.7	+2	42.4	92
23.1	5	44.2	97
25.2	11	46.5	110
26.6	17	48.8	117
33.6	54	49.4	119
38.2	68		

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.048	-0.130	0.483	-0.874
.097	.227	.579	-1.023
.145	.320	.869	.532
.290	.550	.965	.724
.386	.720		

Steele and Johnson, 1904

%	f.t.	%	f.t.
	(7+1)		
3.23	- 0.47	25.48	15
5.24	- 0.72	27.49	22.6
10.69	- 1.46	28.00	22.8
14.70	- 2.15	29.81	30.0
17.47	- 2.85	29.82	30.0
20.47	- 5	30.53	32.3
21.40	0	31.39	33.0
24.00	+ 9	31.27	34.0
	(6+1)I		
30.25	33.0	32.45	44.7
30.35	32.3	33.40	50.0
30.40	35.6	33,63	51.0
30.47	34.0	34.36	53.0
32.42	44.7		
	(6+1)II		
34.43	54.5	37.79	73.0
34.81	57.0	38.72	80.0
35.41	60.0	40.44	89.0
36.86	69.0	43.41	99.0
37.29	70.0		

E : -4.15

tr.t. : (7+1) - (6+1)I 31.5°

(6+1)I - (6+1)II 53.3°

Koppel, 1905

E : ice + (7+1) - 3.9°

Lattey, 1923

%	f.t.
25.45	14.5
32.63	45.6

Tanzov, 1924

f.t.	% (7+1)	% (6+1) blue	% (6+1) green
20	27.39	28.62	30.73
25	28.77	29.37	31.21
30	30.21	30.35	31.78
35	31.98	31.31	32.31
40	33.51	32.25	32.96

Cavon and Johnston, 1927

27.94% f.t.= 25°

Chrétien and Rohmer, 1934

%	f.t.	
0	0	
16.6	-2	
20.7	-3.15	E
21.8	0	(7+1)
27.4	+20	
30.4	29.1	(7+1)-(6+1) I
32.4	40	(6+1)
35.4	60	
36.3	60.3	(6+1)I - (6+1)II
39.4	80	(6+1)II
45.0	100	
46.7	105	
48.2	108	
40.83	84.8	(6+1) - (1+1)
41.01	100	(1+1)
42.23	90.3	(6+1) - (2+1)
42.20	100	(2+1)
42.20	105	
43.82	96.4	(6+1) - (3+1)
43.50	100	(3+1)
43.18	106	
43.98	97.2	(6+1) - (4+1)
44.13	100	(4+1)
44.75	107	
44.19	98	(6+1) - (5+1)
44.44	100	(5+1)
45.95	107.5	

Rohmer, 1939

%	f.t.	%	f.t.
	(7+1)		
9.1	- 0.9	27.7	20
12.7	- 1.6	29.2	25
21.0	- 3.4 E	29.7	27
21.8	0	30.2	28.5
23.3	+ 5	30.6	30
24.9	+10	30.7	30.7 tr.t.
26.3	+15	31.3	32.6
	(6+1)I		
30.1	+27	33.4	45
30.4	28.5	34.6	50
31.1	32.6	35.0	52
31.5	35	35.3	53.8 tr.t.
32.5	40	36.1	58.2
	(6+1)II		
34.7	+50	40.1	80.5
35.5	55	42.4	90.7
36.0	58.2	44.2=	98 tr.t.
36.3	60	45.1	100.8
36.6	61.8	45.3	101.5
37.1	65	46.3	104
37.9	70	47.1	106
38.9	75	48.1	107.8
	(5+1)		
44.3	98.8	44.8	102
44.4	100.0	44.9	104
44.5	100.8	45.2	105
44.7	101.4	45.4	106
	(4+1)		
44.0	97.2[a]	44.0	101.4
44.0	98.8	44.2	104
44.0	100	44.4	106
	(3+1)		
43.8	96.4[b]	43.6	104
43.6	100	43.5	104.5
43.6	101.5	43.5	106
	(2+1)		
42.2	90.3[c]	42,2	102
42.2	100	42.2	104
	(1+1)		
40.9	84.8[d]	40.9	100
40.9	90	40.9	102
40.9	94	40.9	104

a = tr.t. (6+1)II - (4+1)

b = tr.t. (6+1)II - (3+1)

c = tr.t. (6+1)II - (2+1)

d = tr.t. (6+1)II - (1+1)

Benrath, 1941

%	f.t.	%	f.t.
55.2	150	30	205
50	170	25	203
44.2	195	10	205
40	195	1	205
35	203		

E: 20.76% -3.15

tr.t.:(7+1) - (6+1) 30.4% +29.1

(6+1)I = (6+1)II 36.3% 60.3

(6+1)II - (1+1) 40.8% 84.8

Gelbach and Londerback, 1942

28.42% f.t.= 25°

Favre and Valson, 1874

m	d	m	d
	23.5°		
0	0.997	2	1.292
0.5	1.079	2.5	.358
1	.153	3	.421
1.5	.224		

Klein, 1886

N	d
18°	
0	0.7957
0.5	1.0379
1	.0759
2	.1503
3	.2219

Brümmer, 1902

%	d
15°	
0	0.9993
10	1.0925
20	.1950
25	.2518

Cabrera, 1914

%	d	a.10^6	b.10^7
	15°		
0	0.99913	-	-
1.588	1.01636	158.5	52
4.11	.04420	197.7	43
5.84	.06276	213.7	39
9.40	.10425	246.0	32
15.30	.17816	283.0	22
22.80	.28750	320.7	13

$d^t = d^{15}(1-a(t-15)-b(t-15)^2)$

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
14.499	1.1355
28.998	1.2642
14.42	1.1096
22.84	1.2156

Dupèrier, 1924; Cabrera and Dupèrier, 1925

%	d
20°	
12.425	1.1129
44.40	1.4886

Roberts and Adams, 1939

%	d 18°	d 22°
1.917	1.01877	1.01849
3.783	.03870	.03770
7.310	.07761	.07646
13.660	.15329	.15192
18.707	.21913	.21774
25.245	.30970	.30794

Wagner, 1883

%	η (water°=100) 15°	25°	35°	45°
10.62	94.63	73.46	60.12	49.78
18.19	154.93	119.91	99.50	75.70
25.35	298.55	224.93	173.00	152.43

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.048	1.32711	0.483	1.33978
.097	.32869	.579	.34320
.145	.32989	.869	.35018
.290	.33436	.965	.35285
.386	.33726		

Jones, 1904; Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.048	54.60	0.483	31.16
.097	47.10	.579	29.56
.145	44.80	.869	25.08
.290	37.67	.965	23.40
.386	35.92		

Dupèrier, 1924; Cabrera and Dupèrier, 1925

%	χ
20°	
12.425	2.35
44.40	0.25

Gelbach and Londerbach, 1942

M	m	d
	25°	
0.06343	0.06357	1.00758
0.34504	0.34537	1.05250
0.74331	0.74465	1.11332
1.1612	1.1659	1.17592
1.6448	1.6575	1.24685
2.0538	2.0774	1.30546
2.3467	2.3806	1:34889
2.5246	2.5658	1.37460

Brümmer, 1902

%	σ
15°	
0	74.92
10	73.73
20	76.09
25	77.05

Roberts and Adams, 1939

w.l. (Å)	n_D 25.245 %	n_D 18.707 %
	20°	
5769.60	1.38721	1.36113
5460.72	.38849	.36229
4916.04	.39129	.36486
4358.30	.39493	.37578
3663.27	absorption	.37578
3341.47	1.40817	.38084
3131.70	.41279	.38516
2803.50	.42291	.39463
2652.07	.42930	.40066
2482.72	.43846	.40918

Klein, 1886

N	κ 18°	κ 26°	$\tau \cdot 10^4$
0.5	145.5	172.5	231
1	241.3	285.2	227
2	366.7	437.4	241
3	430.0	515.3	250

Water + Potassium nickel sulfate ($K_2NiS_2O_8$)

Tobler, 1855

%	f.t.	%	f.t.	
3.79	0	12.75	36	
6.70	10	16.33	49	
7.90	14	18.42	55	(6+1)
9.21	20	19.68	60	
11.80	30	23.57	75	

von Hauer, 1858

%	f.t.
8.729	20
12.270	40
17.555	60
22.021	80

Water + Nickel ammonium sulfate ($NiN_2H_8O_8S_2$)

Tobler, 1855

%	f.t.	%	f.t.
1.8	3.5	10.3	40
3.1	10	12.6	50
5.5	16	14.3	59
5.6	20	15.8	68
7.7	30	22.2	85

von Hauer, 1858

%	f.t.
9.395	20
13.153	40
18.622	60
23.094	80

Water + Nickel sodium sulfate ($NiNa_2O_8S_2$)

Koppel, 1905

%	f.t.
(4+1)	
29.84	40
31.19	35
32.08	30
34.39	25
35.97	20

Water + Nickel benzenesulfonate ($NiC_{12}H_{10}O_6S_2$)

Ephraim and Seger, 1925

c	f.t.	c	f.t.
12.561	18	26.685	64.5
16.518	34	33.295	80.5
20.972	49.5	34.044	82

Water + Nickel p-phenolsulfonate ($NiC_{12}H_{10}O_8S_2$)

Guerreshi, 1949

%	f.t.	%	f.t.
20.99	15	29.07	34.7
22.36	18.7	31.33	39.8
22.79	20.0	34.21	46.2
24.00	23.0	35.89	50.2
24.80	24.8	37.84	54.8
25.21	25.6	40.31	60.6
27.04	30.0	42.41	65.7

Water + Cobalt chloride ($CoCl_2$)

Heterogeneous equilibria.

Tammann, 1885

%	p	%	p
	100°		
5.70	746.1	22.10	667.7
8.79	736.3	29.32	613.7
13.47	716.4	35.02	567.0
19.38	684.8	36.59	552.5

Lescoeur, 1890

t	p dissoc. (6+1)	p dissoc. (2+1)
10	1.8	-
15	2.3	-
20	4.0	-
25	5.7	-
30	8.05	-
40	14.9	-
100	-	86

Charpy, 1891 (fig.)

t	p		t	p	
	32%				
21	7	red	78	86	blue
28	11	"	80	93	"
35	15	"	82	103	"
40	19	"	86	114	"
52	29		88	120	"
59	39				
70	59				

Derby and Yngve , 1916

t	p sat.sol.	p dissoc.
	(6+1)	(6+1) - (2+1)
23.05	-	6.4
23.40	19.9	-
24.19	14.7	6.8
28.00	-	9.0
28.68	18.4	9.6
31.9	-	12.1
32.65	22.4	12.4
35.98	26.6	16.5
39.47	30.5	20.5
39.87	31.1	21.0
39.96	-	21.5
43.92	36.8	27.8
46.84	41.3	33.3
48.60	-	38.4
49.23	44.8	39.7
50.02	-	42.0
51.15	47.2	46.0
51.58	48.0	46.6
52.13	-	47.7
52.25	48.6	48.6
	(2+1)	
56.44	58.7	
62.12	79.0	
68.26	105.1	
78.87	107.9	

Robinson and Stokes, 1940

m	activity coeff.	m	activity coeff.
	25°		
0.1	0.526	0.9	0.514
.2	.482	1.0	.533
.3	.466	.2	.578
.4	.463	.4	.635
.5	.465	.6	.706
.6	.473	.8	.785
.7	.483	2.0	.884
.8	.496		

Stokes, 1948

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.857	1.2	1.141
.2	.869	.4	.208
.3	.886	.6	.274
.4	.907	.8	.339
.5	.932	2.0	.404
.6	.959	2.5	.564
.7	.982	3.0	.711
.8	1.011	3.5	.821
.9	.043	4.0	.896
1.0	.075		

Rüdorff, 1872

%	f.t.	%	f.t.
4.45	-2.05	15.62	-10.35
9.11	-4.50	18.20	-13.80
12.60	-7.35	20.49	-16.80

van Bemmelen, 1892-1893 (fig.)

%	f.t.	%	f.t.
29	20	43	60
31	30	44	80
35	40	45	100
40	50		

tr.t. (6+1) - (4+1) 52°

Etard, 1894

%	f.t.	%	f.t.
24.7	-22	41.7	45
28.0	-4	46.7	49
31.2	+7	48.4	56
31.3	11	48.8	78
32.5	12	50.2	94
34.4	25	52.3	112
37.5	34	53.5	120
39.8	41		

Jones and Getman, 1904 , Jones, 1904;
Jones and Bassett, 1905

M	f.t.	M	f.t.
0.0699	-0.325	0.800	-4.700
.1279	-0.631	0.900	-5.420
.1918	-0.946	1.000	-6.240
.3197	-1.640	1.500	-11.520
.4475	-2.427	2.000	-19.000
.5114	-2.817	2.500	-27.500
.6393	-3.658	2.760	-33.500
.700	-4.000		

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.01	-0.06241	0.50	-2.8371
.025	-0.1394	0.75	-4.5860
.050	-0.2609	1.00	-6.7157
.075	-0.3356	1.50	-12.1308
.100	-0.5110	2.00	-17.7342
.250	-1.3040		

Bassett and Sanderson, 1932

%	f.t.
29.50	0 (6+1)
35.87	25
40.96	40
49.5	80 (2+1)

Benrath, 1934

%	f.t.	%	f.t.
30.30	0 (6+1)	46.28	50 (4+1)
33.82	17.5	47.64	55
39.69	40.5	48.40	60 (2+1)
42.70	43	49.53	75
44.63	47.5	51.48	99

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987	13	1.1330
1	1.0086	14	.1445
2	.0185	15	.1564
3	.0284	16	.1696
4	.0383	17	.1828
5	.0483	18	.1961
6	.0581	19	.2094
7	.0681	20	.2229
8	.0781	21	.2380
9	.0881	22	.2530
10	.0983	23	.2681
11	.1098	24	.2832
12	.1214	25	.2985

Favre and Valson, 1874

m	d	m	d
	23.9°		
0	0.997	2	1.213
0.5	1.058	2.5	.260
1	.112	3	.304
1.5	.164		

Heydweiller, 1921

N	d
18°	
0.5	1.02901
1	.0578
2	.1138
3	.1684
4	.2221

Miescher, 1930

c	d
20°	
5.48	1.0469
8.28	1.0712
25.7	1.2163

Wagner, 1883

%	η (water°=100) 15°	25°	35°	45°
7.970	83.01	65.07	53.58	44.86
14.858	111.60	85.11	73.69	58.76
22.270	161.60	126.60	101.60	82.58

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.010	113.06	0.50	73.86
0.025	104.86	0.75	66.15
0.050	99.30	1.00	59.58
0.075	96.00	1.50	48.59
0.10	92.26	2.00	38.51
0.25	83.79		

Heydweiller, 1921

N	λ
18°	
0.5	68.6
1	60
2	48.50
3	39.2
4	51.68

Miescher, 1930

w.l.	Verdet's constant . 10^5 c = 5.48	c = 8.28	c = 25.7
	20-21°		
6530	1096	1118	1296
6420	1134	1159	1346
6320	1175	1203	1409
6220	1221	1248	1476
6120	1262	1294	1550
6020	1312	1348	1633
5920	1365	1403	1728
5825	1421	1470	1835
5715	1486	1538	1954
5615	1554	1607	2043
5515	1613	1675	-
5400	1675	1734	-
5310	1728	1776	-
5210	1768	1795	-
5105	1807	-	-
5005	1853	-	-
4905	1910	-	-
4805	1970	-	-
4705	2075	-	-
4605	2200	-	-
4505	2313	-	-
4400	2451	-	-

Teudt, 1900

t	U	t	U
16.40%		8.91%	
31.9	0.7674	31.9	0.8654
33.7	.7790	33.2	.8708
34.3	.7857	33.8	.8711
35.6	.7799	37.3	.8546
36.8	.7821	38.5	.8740
38.9	.7922	40.8	.8830
41.5	.7948	43.7	.8896
43.1	.7994	46.2	.8935
47.6	.7959	49.3	.8977
49.2	.7777	52.7	.8995
51.0	.7949	54.2	.8965
53.4	.7870		

Kapustinskii, Yakuchevskii and Drakin, 1953

%	U	%	U
0	0.9980	13.09	0.8331
4.09	.9425	16.16	.8021
7.55	.8975	17.77	.7713
9.82	.8701	22.63	.7386

Water + Cobalt bromide ($CoBr_2$)

Etard, 1894

%	f.t.
66.7	59
66.8	75
68.1	97

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.04802	73.3
1	.0953	65.5
2	.1893	54.7
3	.2822	45.2
4	.3744	37.56

Water + Cobalt iodide (CoI_2)

Geller, 1911

t	p	t	p	t	p
23.84%		13.53%		6.50%	
16.30	12.208	16.10	13.171	16.6	13.789
18.98	14.278	18.95	15.511	19.7	16.487
22.30	16.932	22.0	18.710	24.05	21.370
25.34	21.838	25.25	22.535	27.8	27.514
27.87	25.628	28.0	27.078	31.31	33.131
30.60	30.154	30.5	31.239	34.1	39.320
33.80	36.438	33.3	36.447	37.4	46.603
36.10	41.829	36.4	43.590	39.65	53.060
38.86	48.510	39.34	51.250	42.3	60.187
41.52	56.354	41.78	58.145	44.4	66.426
44.20	63.719	44.5	66.959	47.1	76.711
46.70	71.788	47.0	75.139	49.23	85.551
48.82	80.024	49.38	84.497		
50.90	87.943				

Etard, 1894

%	f.t.	%	f.t.
52.4	-22	73.0	34
56.7	-8	79.0	46
58.9	-2	79.2	60
61.4	+9	80.7	82
61.6	14	80.9	111
66.4	25	83.1	156

Water + Cobalt thiocyanate ($CoC_2N_2S_2$)

Wernicke, 1912

mol%	d	mol%	d
0.2°		95°	
5.05	1.0052	4.99	0.9695
5.39	.0062	5.30	.9696
6.59	.0073	6.55	.9712
7.77	.0089	7.70	.9727
9.83	.0123	9.76	.9753
28.6	.0460	53.7	1.0937
54.5	.1378	69.0	1.2063
70.0	.2614		

33.3% 25° d= 1.0552 75.2° d=1.0306

m	t	η(water=1)
2	25	1.195
	75	1.1825

Water + Cobalt chlorate ($CoCl_2O_6$)

Meusser, 1902

%	f.t.	%	f.t.
29.97	-12	64.19	+18 (4+1)
37.40	-22	64.39	21
53.30	-21 (6+1)	67.09	35
53.61	-19	69.66	47
57.45	0	76.12	61
61.83	+10.5		

d sat.sol. 18° 1.861

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.04375	61.8
1	.0868	54.3
2	.1717	43.74
3	.2551	35.15
4	.3371	27.91
5	.4183	21.60

Water + Cobalt nitrate (CoN_2O_6)

Tammann, 1885

%	p	%	p
	100°		
13.33	728.7	41.13	552.2
20.52	699.9	46.58	491.9
25.75	677.9	54.83	401.6
34.47	615.7		

Stokes, 1948

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.854	1.4	1.143
.2	.861	.6	.199
.3	.875	.8	.258
.4	.892	2.0	.317
.5	.914	2.5	.468
.6	.936	3.0	.620
.7	.958	3.5	.769
.8	.981	4.0	.913
.9	1.007	4.5	2.053
1.0	1.033	5.0	2.196
1.2	1.087	5.5	2.323

Funk, 1899

%	f.t.	%	f.t.
29.49	-14	49.73	+18
32.85	-18	55.96	41
36.35	-22	62.88	56
39.45	-26 (9+1)		
40.40	-23.5	61.74	55 (3+1)
42.77	-20.5	62.88	62
		64.89	70
41.55	-21 (6+1)	68.84	84
43.69	-10	77.21	91
44.85	-4		
45.66	0		

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	f.t.	M	f.t.
0.0748	-0.352	1.0462	-6.025
.1495	-0.685	.3451	-8.418
.2989	-1.388	.4945	-9.811
.4484	-2.198	2.0000	-17.500
.7473	-3.935	2.5700	-26.500

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.10	-0.0553	0.50	- 2.708
0.025	-0.1341	0.75	- 4.338
0.05	-0.2572	1.00	- 6.220
0.075	-0.3812	1.50	-10.888
0.10	-0.5005	2.00	-18.863
0.25	-1.2705		

Valdman and Klyachko-Gurrvich, 1935

%	f.t.
50.57	25 (6+1)
67.86	80 (3+1)

Franz, 1872

%	d	%	d	%	d
		17.5°			
1	1.0079	15	1.1364	28	1.2911
2	.0171	16	.1476	29	.3041
3	.0263	17	.1586	30	.3173
4	.0355	18	.1698	31	.3344
5	.0449	19	.1810	32	.3455
6	.0537	20	.1920	33	.3595
7	.0626	21	.2040	34	.3736
8	.0715	22	.2176	35	.3878
9	.0804	23	.2280	36	.4031
10	.0892	24	.2400	37	.4184
11	.0986	25	.2521	38	.4337
12	.1080	26	.2651	39	.4489
13	.1174	27	.2781	40	.4643
14	.1268				

Traube, 1895

%	d	%	d
	20°		
0	0.99823	15.327	1.12925
6.364	1.04939	19.618	1.17096
9.758	1.07807	27.319	1.25215

Jones and Getman, 1904; Jones, 1904; Jones and Bassett, 1905

M	d	M	d
	0°		
0.0748	1.011176	1.0462	1.151260
.1495	.022864	.3451	.191612
.2989	.042668	.4945	.211552
.4484	.065092	2.0000	.284688
.7473	.106720	2.5700	.358820

Jones and Pearce, 1907

M	d	M	d
	20°		
0	0.99823	0.25	1.03653
0.10	0.999720	.50	.07225
.025	1.002087	.75	.11007
.05	.005792	1.00	.14410
.075	.009496	1.50	.21504
.10	.013285	2.00	.28348

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.03665	3	1.2144
1	.0729	4	.2836
2	.1444	5	.3516

Miescher, 1930

c	d
20°	
6.38	1.0494
16.20	.1265
55.9	.4221

Wagner, 1883

%	η (water°=100)			
	15°	25°	35°	45°
8.280	74.66	57.95	48.74	39.84
15.960	86.96	69.21	55.38	44.95
24.528	110.40	88.02	71.47	59.06

Jones and Getman, 1904
Jones, 1904 and Jones and Bassett, 1905

M	n_D	M	n_D
	0°		
0.0748	1.32790	0.7473	1.34606
0.1495	.32989	1.0462	.35285
0.2989	.33395	1.3451	.36053
0.4484	.33837	1.4945	.36347

Jones and Getman, 1904
Jones , 1904 and Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.0748	43.84	1.0462	27.30
0.1495	40.90	1.3451	24.08
0.2989	37.50	1.4945	24.73
0.4484	34.73	2.0000	13.04
0.7473	31.21	2.5700	13.29

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.010	108.66	0.50	72.94
0.025	100.69	0.75	65.32
0.050	96.12	1.00	58.89
0.075	91.67	1.50	47.57
0.10	89.94	2.00	37.10
0.25	81.48		

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	66.2	3	37.6
1	58.0	4	29.9
2	46.2	5	23.3

Miescher, 1930

w.l.	Verdet's constant.10^5	w.l.	Verdet's constant.10^5	w.l.	Verdet's constant.10^5
		c=1.0494			
6530	1030				
6530	1030	5715	1385	5005	1748
6420	1062	5615	1445	4905	1799
6320	1104	5515	1504	4805	1875
6220	1143	5410	1564	4705	1956
6120	1185	5310	1611	4605	2067
6020	1229	5210	1657	4505	2192
5920	1276	5105	1702	4400	2317
5815	1326				
		c=1.1265			
6530	991	6120	1153	5715	1379
6420	1034	6020	1199	5615	1449
6320	1066	5920	1252	5515	1512
6220	1110	5815	1310	5410	1568
		c=1.4221			
6530	807	6220	927	5920	1104
6420	846	6120	979	5815	1193
6320	884	6020	1032	5715	1298

Water + Cobalt perchlorate ($CoCl_2O_8$)

Terlikowski, 1911 and Golblum and Terlikowski, 1912

%	f.t.	d
24.60	-10.9	-
29.53	-21.3	-
36.82	-30.7	-
-	-62.2 E	-
45.44	-30.7 (9+1)	-
47.52	-21.3	-
64.02	0	1.5639
65.09	+7.5	.5658
66.24	18	.5670
71.75	26 (5+1)	.5811
73.12	45	.5878

Water + Cobalt sulfate ($CoSO_4$)

Heterogeneous equilibria.

Wiedemann, 1874

t	p dissoc.	t	p dissoc.
25	17.2	55	93.9
30	22.9	60	117.1
35	33.2	65	145.6
40	43.6	70	177.1
45	56.7	75	213.2
50	61.4	80	255.4

Tammann, 1885

t	p 0%	p 21.16%	p 30.56%
24.7	25.65	24.0	22.8
39.4	33.77	-	37.1
61.9	42.12	59.6	56.4
79.4	46.94	76.4	73.2
107.6	53.07	104.2	99.7
154.0	60.60	150.2	143.6
188.1	65.06	182.3	174.9
220.5	68.65	213.5	203.8
261.6	72.62	253.7	243.0
296.8	75.61	288.4	277.3
344.1	79.20	333.6	321.7
399.6	82.92	388.1	373.8
459.1	86.46	444.9	428.8
528.6	90.13	512.9	494.6
604.9	93.73	587.5	567.9
682.4	97.01	665.6	644.2
763.1	100.11	743.4	720.7

%	p	%	p
	100°		
9.53	752.5	26.65	731.9
16.24	746.7	34.28	705.8
22.41	739.6		

Lescoeur, 1895

t	p sat.sol.
17	11.9
20	13.7
60	173
80	295

Kahlenberg, 1901

%	b.t.	%	b.t.
	760 mm		
2.05	100.07	17.06	100.58
4.25	100.11	18.20	100.64
8.04	100.21	19.40	100.71
8.76	100.27	20.72	100.79
11.45	100.35	22.33	100.90
13.89	100.46	24.72	101.08

Rohmer, 1934 (fig.)

sat.sol. b.t.= 104.5

Tobler, 1855

%	f.t.	%	f.t.
20.8	3	31.6	35
23.4	10	33.6	44
26.7	20	35.6	50
28.0	24	37.2	60
28.6	29	39.7	70

Mulder, 1864

%	f.t.	%	f.t.
19.74	0	35.10	54
23.67	14.25	41.42	82.5
25.98	20.50	46.49	106.4 b.t.
31.69	40.25		

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.0567	-0.143	0.6799	-1.187
.1133	-0.245	0.9066	-1.614
.2267	-0.435	1.1333	-2.073
.3399	-0.625		

Koppel, 1905

%	f.t.	%	f.t.
		(7+1)	
20.34	0	28.24	25
21.90	5	29.70	30
23.40	10	31.40	35
24.83	15	32.81	40
26.58	20		

E : -3° tr.t. : (7+1) - (6+1) +40.7

Rohmer, 1934 (fig.)

%	f.t.	
0	0	
9.1	-1.2	
16.7	-2.3	
19.0	-2.7	E
23.1	+10	(7+1)
28.6	30	
32.9	43.3	(7+1) - (6+1)
36.1	60.0	(6+1)
36.7	64.2	(6+1) - (1+1)
33.3	80	(1+1)
28.2	95	
38.8	72.9	(6+1) - (2+1)
38.1	100	(2+1)
40.5	80.4	(6+1) - (4+1)
43.2	100	(4+1)

Benrath, 1941

%	f.t.	%	f.t.
47	115	30	185
45	135	20	205
41.5	153	10	205
37.2	170	5	205

Properties of phases.

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
12.217	1.1131
24.433	1.2218
12.356	1.1139
24.712	1.2238

Roberts, 1930

%	d		%	d	
	20°	25°		20°	25°
4.33	1.04249	1.04112	19.14	1.21308	1.21101
6.45	.06524	.06362	22.26	.25525	.25313
10.23	.10664	.10500	25.09	-	.29465

Miescher, 1930

c	d
20°	
5.8	1.0545
6.29	.0596
14.85	.1390
29.8	.2720

Cantelo and Phifer, 1933

M	molar volume	M	molar volume
	25°		
2.250	16.82	1.121	11.27
1.954	16.29	0.9721	10.23
.827	15.73	.8752	8.976
.731	15.12	.7261	8.264
.615	14.47	.6473	7.473
.511	13.76	.5301	6.498
.389	13.01	.3587	5.271
.265	12.18		

Wagner, 1883

%	η(water°=100)			
	15°	25°	35°	45°
7.239	86.72	68.72	54.96	45.08
14.156	117.81	95.51	76.03	61.66
21.167	193.60	146.20	113.00	89.95

Roberts, 1930

w.l. (Å)	n					
	0%	4.33%	6.45%	10.23%	19.14%	22.26%
			20°			
7065	1.33003	1.33796	-	-	1.36708	1.37400
6678	.33087	.33882	1.34284	1.35016	.36802	.37498
6563	.33115	.33913	.34312	-	.36834	-
5876	.33305	.34106	.34509	1.35241	.37044	1.37743
5461	.33440	.34254	.34666	.35382	-	-
4471	.33945	.34762	.35173	-	-	-
4358	.34027	.34845	.35258	-	1.37841	-
4047	.34284	.35108	.35518	1.36254	.38128	1.38839
3886	.34432	.35263	.35678	.36422	.38291	.39013
3341	.35165	.36015	.36436	.37193	.39106	.39834
3187	-	.36317	.36745	-	.39426	.40169
3131	1.35567	.36435	.36860	1.37632	.39556	.40293
2893	.36168	.37050	.37478	.38254	.40220	.40968
2801	.36442	.37334	.37779	.38543	-	-
2576	.37338	.38247	.38698	.39482	1.41506	1.42276
2482	.37809	.38732	-	.39977	-	.42803
2373	.38434	.39366	1.39827	.40627	1.42705	.43496

t	w.l.(Å)	(α)magn.
	0%	
18.70	6780	5.537
"	6104	6.760
"	5893	7.265
"	5780	7.519
"	5461	8.464
"	4916	10.669
	6.45%	
18.70	6708	5.447
"	6104	6.698
"	5893	7.252
"	5780	7.658
"	5461	8.748
"	4916	10.09
"	5194	9.40
"	4529	12.02
"	4384	13.05
	19.4%	
15.5	6908	4.86
18.9	6708	5.25
17.0	6234	6.24
17.0	6104	6.57
15.1	5893	7.27
15.5	5780	7.746
	5461	9.13

t	w.l.(Å)	(α)magn.
	4.33%	
18.8	6908	5.01
19.0	6708	5.46
19.0	6104	6.736
18.5	5893	7.289
16.9	5780	7.548
17.9	5461	8.594
19.5	4916	10.18
"	5230	9.36
"	5167	9.515
"	4958	10.04
	10.23%	
18.8	6908	5.07
16.9	6708	5.410
16.6	6104	6.70
16.8	5893	7.208
17.2	5780	7.610
19.8	5461	8.75
16.9	568.2	7.94
	22.26%	
19.2	6708	5.29
19.2	6104	6.48
19.3	5893	7.281
19.3	5780	7.739
	5461	absorpt.

Jones, 1904; Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.0567	57.49	0.6799	30.67
.1133	48.56	0.9066	27.16
.2267	42.06	1.1333	23.15
.3399	38.07		

Miescher, 1930

w.l. Å	0%	Verdet's constant .10[5] 5.8g/100cc	6.92g/100cc	14.85g/100cc	29.8g/100cc
			20-21°		
6530	1058	1052	1048	1034	1001
6420	1095	1082	1086	1068	1038
6320	1135	1118	1122	1114	1078
6220	1171	1163	1169	1155	1128
6120	1209	1201	1207	1205	1173
6020	1254	1248	1254	1248	1229
5920	1298	1298	1300	1302	1294
5815	1344	1352	1359	1363	1369
5715	1394	1411	1417	1433	1462
5615	1451	1474	1482	1510	1562
5515	1510	1532	1536	1577	1621
5410	1569	1589	1597	1625	-
5310	1637	1641	1643	1647	-
5210	1703	1684	1684	-	-
5105	1781	1720	1724	-	-
5005	1858	1772	1760	-	-

Water + Cobalt chloride pentammine ($CoCl_3H_{15}N_5$)

Kurnakow, 1895

%	f.t.
14.09	0
19.73	16.2
19.92	16.9

Water + Cobalt ammonium sulfate ($CoN_2H_8O_8S_2$)

Tobler, 1855

%	f.t.	%	f.t.
7.4	0	18.2	40
10.4	10	20.0	45
13.1	18	22.3	50
14.6	23	26.2	60
16.4	35	30.2	75

von Hauer, 1858

%	f.t.
14.927	20
20.782	40
25.579	60
32.988	80

Water + Cobalt sodium sulfate ($CoNa_2O_8S_2$)

Koppel, 1905

%	f.t.	
34.01	20	(4+1)
32.75	25	
30.91	30	
30.26	35	
29.09	40	

Water + Potassium cobalt sulfate ($K_2CoS_2O_8$)

Tobler, 1855

%	f.t.	%	f.t.
12.07	0	25.72	30
17.37	12	27.46	35
18.47	15	29.55	40
21.28	20	34.55	49
23.47	25		

(6+1)

von Hauer, 1858

%	f.t.
13.968	20
19.539	40
24.372	60
31.816	80

(4+1)

Water + Cobalt benzenesulfonate ($CoC_{12}H_{10}O_6S_2$)

Ephraim and Seger, 1925

c	f.t.
7.575	17
11.086	36
14.144	50
19.717	65
26.735	80.5

Water + Tris-Ethylenediamine cobalt chloride ($C_6H_{24}N_6Cl_3Co$)

Rostkovsky, 1927

%	f.t.	%	f.t.
5	-0.7	44.8	55
10	-1.8	47.5	60
14.7	0	49.8	65
18.1	+5	52.0	70
21.3	10	53.8	75
24.1	15	55.6	80
26.6	20.3	57.1	85
29.2	25	58.4	90
32.0	30	59.7	95
34.8	35	61.1	100
37.5	40	62.3	105
39.8	45	63.9	110
42.3	50 (3+1)		

LIV. WATER + SALTS OF OTHER POLYVALENT METALS .

Water + Lithium cyanoplatinite ($Li_2PtC_4N_4$)

Terrey and Jolly, 1923

%	f.t.	%	f.t.
105.0	0.0	154.5	39.2
139.5	16.3	160.3	40.0
141.5	22.2	160.5	42.5
153.5	23.0	181.2	42.7
144.8	24.1	188.2	43.1
144.7	25.0	162.3	43.2
153.0	25.7	196.0	45.0
146.0	25.8	165.0	45.7
147.5	26.5	176.1	46.6
151.2	26.8	186.0	47.2
148.5	28.5	173.0	48.0
152.3	30.1	173.1	49.2
156.1	30.5	175.8	50.1
152.0	31.5	182.0	53.0
152.4	32.2	173.8	55.2
166.6	34.9	178.0	60.7
154.6	35.0	185.7	64.8
155.6	35.8	205.2	66.0
173.0	37.3	204.0	71.0
155.3	38.0	213.7	78.2
158.2	38.8	229.10	88.2
		238.7	89.8

Water + Potassium cyanoplatinite ($K_2PtC_4N_4$)

Terrey and Jolly, 1923

%	f.t.	%	f.t.
10.40	0.10 (5+1)	56.05	55.40 (2+1)
16.50	9.80	58.18	60.40
20.96	14.40	61.08	67.50
22.26	16.38 (3+1)	63.67	74.50 (1+1)
22.49	17.42	63.41	78.20
24.71	20.05	64.10	83.60
27.45	22.65	64.79	87.20
29.55	25.00	67.84	95.00
39.08	35.00		
43.90	39.75		
48.80	45.00		
52.10	49.90		

Water + Cobalt benzenesulfonate ($CoC_{12}H_{10}O_6S_2$)

Ephraim and Seger, 1925

c	f.t.
7.575	17
11.086	36
14.144	50
19.717	65
26.735	80.5

Water + Tris-Ethylenediamine cobalt chloride ($C_6H_{24}N_6Cl_3Co$)

Rostkovsky, 1927

%	f.t.	%	f.t.
5	-0.7	44.8	55
10	-1.8	47.5	60
14.7	0	49.8	65
18.1	+5	52.0	70
21.3	10	53.8	75
24.1	15	55.6	80
26.6	20.3	57.1	85
29.2	25	58.4	90
32.0	30	59.7	95
34.8	35	61.1	100
37.5	40	62.3	105
39.8	45	63.9	110
42.3	50 (3+1)		

LIV. WATER + SALTS OF OTHER POLYVALENT METALS .

Water + Lithium cyanoplatinite ($Li_2PtC_4N_4$)

Terrey and Jolly, 1923

%	f.t.	%	f.t.
105.0	0.0	154.5	39.2
139.5	16.3	160.3	40.0
141.5	22.2	160.5	42.5
153.5	23.0	181.2	42.7
144.8	24.1	188.2	43.1
144.7	25.0	162.3	43.2
153.0	25.7	196.0	45.0
146.0	25.8	165.0	45.7
147.5	26.5	176.1	46.6
151.2	26.8	186.0	47.2
148.5	28.5	173.0	48.0
152.3	30.1	173.1	49.2
156.1	30.5	175.8	50.1
152.0	31.5	182.0	53.0
152.4	32.2	173.8	55.2
166.6	34.9	178.0	60.7
154.6	35.0	185.7	64.8
155.6	35.8	205.2	66.0
173.0	37.3	204.0	71.0
155.3	38.0	213.7	78.2
158.2	38.8	229.10	88.2
		238.7	89.8

Water + Potassium cyanoplatinite ($K_2PtC_4N_4$)

Terrey and Jolly, 1923

%	f.t.	%	f.t.
10.40	0.10 (5+1)	56.05	55.40 (2+1)
16.50	9.80	58.18	60.40
20.96	14.40	61.08	67.50
22.26	16.38 (3+1)	63.67	74.50 (1+1)
22.49	17.42	63.41	78.20
24.71	20.05	64.10	83.60
27.45	22.65	64.79	87.20
29.55	25.00	67.84	95.00
39.08	35.00		
43.90	39.75		
48.80	45.00		
52.10	49.90		

Water + Lithium chloraurate ($LiAuCl_4$)

Rosenbladt, 1886

%	f.t.	%	f.t.
53.1	10	72.0	50
57.7	20	76.4	60
62.5	30	81.0	70
67.3	40	85.7	80

Water + Sodium chloraurate ($NaAuCl_4$)

Rosenbladt, 1886

%	f.t.	%	f.t.
58.2	10	69.4	40
60.2	20	77.5	50
64.0	30	90.0	60

Water + Potassium chloraurate ($KAuCl_4$)

Rosenbladt, 1886

%	f.t.	%	f.t.
27.7	10	59.2	40
38.2	20	70.0	50
48.7	30	80.2	60

Water + Rubidium chloraurate ($RbAuCl_4$)

Rosenbladt, 1886

%	f.t.	%	f.t.
4.6	10	26.6	60
9.0	20	31.0	70
13.4	30	35.3	80
17.7	40	39.7	90
22.2	50	44.2	100

Water + Cesium chloraurate ($CsAuCl_4$)

Rosenbladt, 1886

%	f.t.	%	f.t.
0.5	10	8.2	60
0.8	20	12.0	70
1.7	30	16.3	80
3.2	40	21.7	90
5.4	50	27.5	100

Water + Aluminum chloride ($AlCl_3$)

Tammann, 1885

%	p	%	p
	100°		
6.02	738.6	23.22	543.9
7.75	729.6	26.94	475.9
13.29	687.5	30.49	411.8
16.43	651.0	31.39	393.6
19.40	609.5		

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.046	-0.276	0.657	-5.120
.076	-0.446	0.876	-7.970
.102	-0.578	1.195	-13.610
.200	-1.148	1.434	-19.518
.299	-1.840	1.593	-23.870
.398	-2.596	2.124	-45.000
.531	-3.830		

Jones and Pearce, 1907

M	f.t.	M	f.t.
0.01	-0.712	0.25	-1.6604
.025	-0.1623	.50	-3.9446
.05	-0.3053	.75	-7.1339
.075	-0.4511	1.00	-11.795
.10	-0.4511(?)	1.50	-25.5
		2.00	-48.5

Jones and Stine, 1908

M	%	f.t.
0.2	2.604	- 1.279
0.4	5.103	- 2.910
0.6	7.494	- 5.144
0.8	9.808	- 7.962
1.0	12.03	-11.795
1.2	14.15	-16.385
1.3617	15.82	-21.75
1.415	-	-23.50

Malquori, 1928

%	f.t.	E	%	f.t.
			(6+1)	
0	0	-	31.03	0
5.25	-5.9	-55.4	31.36	20
9.11	-7.7	-55	31.63	40
12.59	-10.5	-54.3	31.73	60
20.14	-33.9	-54	32.32	80
24.12	-49.2	-55.8		
24.67	-52.0	-54.2		
24.85	-52.9	-55.8		
25.25	-54.7	-55.8		
26.79	-42.5	-56.7		
27.27	-30.1	-55.8		
28.00	-27.3	-53.9		

Tanaka, 1930

%	f.t.	%	f.t.
	(6+1)		
30.48	0	32.17	45
30.82	5	32.32	65
31.66	10	33.23	98
31.96	30		

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.819	0.8	1.220
.2	.841	0.9	.299
.3	.889	1.0	.382
.4	.947	.2	.560
.5	1.008	.4	.749
.6	.074	.6	.951
.7	.145	.8	2.175

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0165	16	1.1487
4	.0346	18	.1692
6	.0528	20	.1905
8	.0713	24	.2352
10	.0900	28	.2834
12	.1091	30	.3089
14	.1287		

Gerlach, 1859

%	d	%	d
	15°		
0	0.99913	21	1.16130
1	1.00634	22	.16990
2	.01355	23	.17850
3	.02075	24	.18711
4	.02795	25	.19571
5	.03516	26	.20479
6	.04262	27	.21387
7	.05007	28	.22299
8	.05753	29	.23202
9	.06498	30	.24110
10	.07243	31	.25074
11	.08025	32	.26039
12	.08807	33	.27004
13	.09588	34	.27968
14	.10370	35	.28933
15	.11151	36	.29952
16	.11975	37	.30971
17	.12799	38	.31990
18	.13622	39	.33009
19	.14445	40	.34029
20	.15269	41	.35105

Jones, 1904; Jones and Getman, 1904

M	d	M	d
	0°		
0	0.999868	0.6372	1.065632
0.2124	1.022372	1.0620	.107808
0.4248	1.044864	1.4868	.146324
		2.1240	.203156

Chéneveau, 1907

%	d	%	d
	22°		
0	0.9978	11.59	1.1068
4.12	1.0371	14.98	.1408
7.98	.0722	18.17	.1738
9.81	.0895		

Jones and Pearce, 1907

M	d	M	d
	20°		
0	0.99823	0.25	1.02729
0.01	.99927	.50	.05519
.025	1.00105	.75	.08239
.05	.00411	1.00	.10857
.075	.00672	1.50	.16102
.10	.00979	2.00	.21163

Chéneveau, 1907

%	t	d
13.86	18.8	1.1343
14.40	18.7	1.1365

Jones and Stine, 1908

%	d	%	d
	0°		
2.604	1.024960	12.03	1.099404
5.103	.046142	14.15	.131508
7.494	.068536	15.82	.149008
9.808	.088612		

Getman and Wilson, 1908

%	d	%	d
	18°		
0	0.9986	20	1.1537
5	1.0360	25	.1967
10	.0733	30	.2421
15	.1124		

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
8.3146	1.0672
14.349	1.1150

Kohner, 1928

equiv./1000g water	d	equiv./1000g water	d
	25°		
0	0.99707	3.85492	1.13340
0.99692	1.03524	5.98371	1.19965
2.00956	1.07138	7.60341	1.24715

Okazaki, 1935

%	d	%	d
	28°		
7.89	1.0616	21.35	1.1838
15.02	.1243	25.13	.2252
18.28	.1560	26.86	.2435
18.72	.1601	30.93	.2882

Guillaume, 1946

%	d
20°	
7.64	1.0696
14.4	.1350
22.3	.2204

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0	1.33395	1.0620	1.36145
.2124	.33297	1.4868	.37493
.4248	.34064	2.1240	.39441
.6372	.34736		

Chéneveau, 1907

%	n_D	%	n_D
	22°		
0	1.3328	11.59	1.3662
4.12	.3442	14.98	.3768
7.98	.3553	18.17	.3871
9.81	.3607		

%	n C	n D	n Tl	n F	n G'
		18.8°			
13.86	1.37350	1.37574	1.37854	1.38083	1.38485
14.40	.37447	.37671	.37912	.38188	.38833

Getman and Wilson, 1908

%	n_D	%	n_D
	18°		
0	1.33314	20	1.39063
5	.34625	25	.40519
10	.35963	30	.42444
15	.37450		

Kohner, 1928

equiv./1000g water	n_D	equiv./1000g water	n_D
	25°		
0	1.33254	3.85492	1.37629
0.99692	1.34488	5.98371	1.39719
2.00956	1.35649	7.60341	1.41193

Okazaki, 1935

%	Verdet's constant . 10^5 (3441 Å)
28°	
7.89	5095
15.02	5683
18.28	6019
18.72	6103
21.35	6275
25.13	6676
26.86	6893
30.93	7210

Mason, Gray and Ernst, 1936

M	Verdet's constant . 10^5					
	10°	20°	25°*	30°	40°	50°
			5893 Å			
0.00	1319	1312	1311	1310	1310	1310
.01	-	1313	1314	1314	1314	1315
.05	1322	1325	1327	1329	1332	1336
.10	1339	1339	1339	1339	1339	1339
.50	1466	1466	1466	1466	1466	1466
1.00	1594	1594	1594	1594	1594	1594
1.50	1721	1721	1721	1721	1721	1721
2.00	1841	1841	1835	1831	1828	1828
2.78	2015	2002	2002	1995	1995	1990
			5460.7 Å			
0.00	1540	1540	1540	1540	1540	1533
.01	-	1547	1547	1547	1547	1547
.05	-	1583	1578	1573	1562	1551
.10	1594	1594	1594	1594	1594	1594
.50	1714	1714	1714	1714	1714	1714
1.00	1861	1861	1861	1861	1861	1861
1.50	2009	2009	2009	2009	2009	2009
2.00	2149	2149	2156	2156	2163	2163
2.78	2377	2377	2370	2364	2350	2350

Guillaume, 1946

%	$n_{5780\ Å}$	$(\alpha)^*$ magn. 10^6
	20°	
7.64	1.3562	4.239
14.4	.3767	.431
22.3	.4035	.648

*in radians, gauss, centim.

Jones, 1904 and Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.0532	124.6	0.69	71.7
0.213	104.6	1.064	59.3
0.425	88.64	1.423	43.1

Jones and Pearce, 1907

M	molecular conductivity	M	molecular conductivity
	0°		
0.010	155.48	0.50	88.60
0.025	141.24	0.75	75.02
0.050	130.44	1.00	61.93
0.075	126.66	1.50	41.62
0.10	122.07	2.00	25.47
0.25	106.90		

Jones and Stine, 1908

M	%	$\varkappa$
	0°	
0.2	2.604	239.83
0.4	5.103	405.30
0.6	7.494	531.15
0.8	9.808	611.80
1.0	12.03	994.04
1.2	14.15	1315.08
1.3617	15.82	1490.08

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	65	3	34.70
1	56.20	4	27.16
2	44.18		

Jauch, 1921

N	U	N	U
	18°		
0.5	0.9593	3	0.8098
1	.9257	4	.7630
2	.8632		

Kapustinskii and Ruzavin, 1955

heat conductivity coefficient c . 10^6

%	c	%	c
	25°		
3.54	1425	12.1	1349
8.43	1388	17.2	1296
10.63	1368	27.1	1206
5.42	1410	17.7	1304
10.8	1367	27.0	1197

Water + Aluminum bromide ($AlBr_3$)

Traube, 1895

%	d
20°	
0	0.99823
6.066	1.04578
10.253	.08136
15.956	.13318
20.896	.18157

Water + Aluminum chlorate ($AlCl_3O_9$)

Roth, 1903

%	f.t.	%	f.t.
0.3581	-0.0775	5.146	-1.1580
0.7416	-0.1535	10.383	-2.6965
2.4411	-0.5105	16.461	-5.4985

Water + Aluminum bromate ($AlBr_3O_9$)

Dobroserdow, 1926

$M_{(9+1)}$	t	n_D	$M_{(9+1)}$	t	n_D
0.0078	17.90	1.33405	0.2094	20.10	1.35724
.0199	20.50	.33484	.2601	19.80	.36393
.0287	20.60	.33586	.3138	19.85	.37179
.0448	20.65	.33736	.4101	19.80	.38693
.0576	20.45	.33895	.4956	19.60	.40331
.0844	20.25	.34193	.6195	19.25	.42637
.1491	20.20	.34874	.6575	19.00	.43477
.1739	20.05	.35201	.7634	18.65	.46076

Water + Aluminum iodate (AlI_3O_6)

Dobroserdow, 1926

$M_{(9+1)}$	t	n_D	$M_{(9+1)}$	t	n_D
0.1054	25.1	1.34365	0.0334	25.7	1.33585
.0750	24.3	.34065	.0175	25.5	.33429
.0645	24.45	.33943	.0082	25.0	.33328
.0548	24.6	.33831	.0066	25.2	.33298
.0483	24.4	.33767	.0035	25.3	.33258
.0352	25.7	.33625	.0023	25.2	.33244

Water + Aluminum nitrate (AlN_3O_9)

Pearce, 1936

M	p	M	p
25°			
0	23.752	1.0	21.011
0.1	23.648	1.5	20.386
.2	23.500	2.0	18.561
.4	23.235	2.5	16.678
.6	22.860	3.0	14.860
.8	22.405	3.1607	14.370
		satd.	

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.0533	-0.333	0.5330	-4.260
.1066	-0.652	.7462	-6.760
.2132	-1.410	.8528	-8.190
.3198	-2.290	1.0660	-11.790

Malquori, 1927

%	f.t.	E	%	f.t.
6.35	-2.4	-27	37.81	0 (9+1)
12.41	-5.8	"	40.15	10
21.35	-13	"	42.99	20
26.20	-18.5	"	44.94	30
28.61	-23	"	46.25	40
30.45	-27	"	49.14	50
32.40	-23	"	50.95	60
33.60	-16	"	54.60	70
34.92	-11.5	"	60.55	90 (8+1)
35.75	-7.3	-26	60.81	95.8
36.81	-3	-26.4	61.12	99.5
			61.44	100
			61.64	103
			62.19	111.5 (6+1)
			62.32	116
			63.01	122
			63.78	129

Tanaka, 1930

%	f.t.
42.23	0 (9+1)
32.11	5
35.75	10
42.64	30
49.82	65

Jones, 1904; Jones and Getman, 1904

M	d	M	d
	0°		
0.0533	1.006460	0.5330	1.084212
.1066	.015432	.7462	.117200
.2132	.033040	.8528	.133132
.3198	.049948	1.0660	.164320

Heydweiller, 1921

N	d	N	d
	18°		
0.5	1.02765	4	1.2117
1	.0546	5	.2620
2	.1079	6	.3118
3	.1603		

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
10.22	1.0703
20.44	1.1414
11.846	1.10822
21.771	1.1502

Pearce and Blackman, 1935

M	d	M	d
	25°		
0.0	0.997071	1.0	1.143200
.1	1.012725	1.5	.206298
.2	.028301	2.0	.262843
.4	.058649	2.5	.318080
.6	.087913	3.0	.357277
.8	.116096	3.1607 satd.	.370250

Guillaume, 1946

%	d
20°	
12	1.1003
20	.1764
27.4	.2534

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.0533	1.32727	0.5330	1.34493
.1066	.32949	.7462	.35240
.2132	.33354	.8528	.35688
.3198	.33726	1.0660	.36283

Guillaume, 1946

%	n 5780 Å	$(\alpha)^*_{magn.}10^6$
	20°	
12	1.3561	3.586
20	.3727	.317
27.4	.3897	.082

*in radians, gauss, centim.

Jones, 1904 and Jones and Getman, 1904

M	molecular conductivity	M	molecular conductivity
	0°		
0.0533	137.85	0.5330	86.10
0.1066	122.00	0.7462	73.27
0.2132	111.89	0.8528	65.98
0.3198	101.60	1.0660	55.21

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	61.8	3	31.76
1	53.3	4	24.09
2	41.07	5	18.16
		6	13

Jauch, 1921

N	U	N	U
	18°		
0.5	0.9563	3	0.8008
1	.9127	4	.7504
2	.8526	5	.7140

Water + Aluminum perchlorate ($AlCl_3O_{12}$)

Geffcken, 1929

equiv./1000g water	d	n_{He}
	25°	
0	0.99707	1.332590
0.2883	1.06167	.3414434
0.3378	.07230	.342883
0.6083	.12846	.350440
0.9925	.20270	.360346
1.4627	.28570	.371324
1.4665	.28638	.371422
1.9925	.37057	.382510
2.1006	.38689	.384671
2.8115	.48678	.397881

Water + Aluminum sulfate ($Al_2S_3O_{12}$)

Heterogeneous equilibria .

Tammann, 1885

t	p 0%	8.74%	14.80%	24.08%	30.00%
43.52	66.6	65.7	65.7	62.5	61.0
47.97	83.6	82.5	82.2	78.8	76.3
51.43	99.3	97.8	97.6	93.2	90.2
56.42	126.4	124.7	124.1	119.0	115.0
60.04	149.7	147.8	147.1	140.8	136.0
63.50	175.4	173.6	172.6	165.8	159.3
67.40	208.7	206.6	205.2	196.5	189.2
69.70	230.8	228.4	227.2	218.3	210.8
74.78	286.7	284.6	282.8	272.9	263.8
76.10	302.9	300.7	298.6	288.0	278.2
79.40	346.9	344.7	342.7	330.4	319.0
81.15	372.4	370.1	368.2	356.1	343.2
84.32	422.3	418.7	416.5	402.4	389.1
89.06	507.6	503.0	-	482.8	466.5
92.00	567.2	562.4	558.4	541.5	524.3
95.04	634.6	629.8	624.9	606.9	588.4
97.28	689.0	684.8	678.6	661.2	639.1
100.21	765.7	759.5	754.0	734.5	710.4

%	p	%	p
	100°		
9.02	752.8	26.70	719.9
14.01	750.6	28.70	708.5
19.22	740.0		

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
0.1	0.420	0.6	0.545
0.2	0.390	0.7	0.625
0.3	0.391	0.8	0.718
0.4	0.421	0.9	0.809
0.5	0.477	1.0	0.922

Kremann and Hüttinger, 1909

%	p dissoc.	%	p dissoc.
	20°		
51.34	17.5	70.37	0.525
54.27	8.05	82.61	.525
57.59	6.825	90.48	.175
61.29	0.525	95.00	.175

Poggiale, 1843

%	f.t.	%	f.t.
23.84	0	37.15	60
25.09	10	39.84	70
26.55	20	42.25	80
28.75	30	44.69	90
31.38	40	47.12	100
34.27	50		

Gerlach, 1889

25.705% f.t.= 15°

Kremann and Hüttinger, 1909

%	f.t.	%	f.t.
8.09	-1.02	17.50	-2.65
10.74	-1.43	19.21	-2.85
14.43	-2.04	24.76	+7.73

Henry and King, 1949

%	f.t.	%	f.t.
28.3	25	30.6	50
28.6	30	31.2	55
28.9	35	31.8	60
29.6	40	32.9	65
30.0	45		

Properties of phases

Bischof, 1850

%	d	%	d
	18.75°		
2	1.0187	16	1.1763
4	.0396	18	.2010
6	.0609	20	.2263
8	.0829	24	.2785
10	.1054	28	.3331
12	.1286	28.45	.3395
14	.1521		

Gerlach, 1889

%	d	%	d
	15°		
0	0.9991	15.423	1.1700
5.141	1.0526	20.564	.2344
10.282	.1095	25.705	.34

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
17.688	1.1558
27.875	1.2381

Kohner, 1928

equiv./1000g water	d
25°	
0	0.99707
0.66737	1.10647
0.66765	1.10653
0.89592	1.14134
1.33354	1.20616
1.63215	1.24789
2.20574	1.32220

Silva and Chenevey, 1945

t	d 4.93%	9.85%	14.79%	19.77%	24.6%
34.9	1.0440	1.0980	1.1539	1.2147	1.2785
45.0	.0399	.0930	.1484	.2092	.2726
55.0	.0347	.0879	.1426	.2033	.2666
65.2	.0288	.0805	.1361	.1964	.2603
75.2	.0222	.0742	.1281	.1881	.2513
85.4	.0196	.0665	.1213	.1811	.2446
95.0	.0188	.0598	.1206	.1751	.2381
	34.5%	39.8%	44.3%		
65.2	1.3968	-	-		
75.2	.3880	-	-		
85.4	.3809	-	-		
95.0	-	1.4542	-		
106.2	-	1.4332	1.5014		

%	d 75.2°	85.4°	%	d 75.2°	85.4°
4.93	1.0222	1.0296	19.77	1.1881	1.1811
9.85	.0742	.0665	24.6	.2513	.2446
14.79	.1281	.1213	34.5	.3880	.3809

Cupples, 1946

M	m	d 15°	25°
0.0286	0.0287	1.0090	1.0068
.0584	.0585	.0191	.0168
.1217	.1218	.0404	.0378
.1859	.1863	.0617	.0588
.2510	.2517	.0829	.0798
.3227	.3242	.1060	.1025
.4721	.4761	.1530	.1490
.627	.636	.2000	.1958
.810	.829	.2547	.2503
1.027	1.063	.3172	.3123

Guillaume, 1946

%	d
20°	
8.6	1.0919
19.4	1:2220
24.9	1.2953

Kohner, 1928

equiv./1000g water	n_D	equiv./1000g water	n_D
	25°		
0	1.33254	1.33354	1.37361
0.66737	.35456	1.63215	.38134
0.66765	.35458	2.20574	.39455
0.89592	.36141		

Guillaume, 1946

%	n 5780 Å	(α)* magn. 10^6
	20°	
8.6	1.3536	3.732
19.4	.3786	.413
24.9	.3911	.250

*in radians, gauss, centim.

Cupples, 1946

M	σ	M	σ
		25°	
0.0286	71.92	0.3227	73.67
.0584	72.04	.4721	74.45
.1217	72.33	.6270	75.56
.1859	72.96	.8100	77.42
.2510	73.29	1.0270	80.78

Johnston, 1907

N	λ	N	λ
	99.4°		
4	19.1	0.25	40.7
2	23.4	0.10	64.4
1	35.7	0.01	171.9
0.5	41.9	0.001	190.8

Marignac, 1876

%	U
21-53°	
3.08	0.9722
5.97	.9465
11.26	.9041
20.26	.8400

Water + Ammonium alum ($AlH_4NO_8S_2$)

Tammann, 1885

%	p
100°	
9.89	749.1
17.53	738.7
25.60	723.1
32.65	700.3
40.44	663.6

Poggiale, 1843

%	f.t.	%	f.t.
2.55	0	17.41	60
4.30	10	21.23	70
6.16	20	26.03	80
8.30	30	33.46	90
10.99	40	41.46	100
13.71	50		

Lescoeur and Mathurin, 1888

t	p dissoc. (24+1)	(6+1)
60	46	-
69.3	82	-
78.8	147	-
88.4	320	-
100	-	75

Bindel, 1890

%	d	%	d
		20°	
5.85	1.0570	20.21	1.2010
10.52	.0981	29.15	.3012
12.52	.1184	37.44	.3898
15.46	.1494		

Kasankine, 1891

d	capillary rise	d	capillary rise
16-17°		32-34°	
1.0464	25.62	1.0579	25.02

Bindel, 1890

%	U	Q diss.
	20°	
5.85	0.942	- 7994
10.52	0.901	- 8126
12.52	0.883	- 8202
15.46	0.858	- 8277
20.21	0.821	- 8472
29.15	0.752	- 9083
37.44	0.691	-10627

Water + Sodium aluminum alum ($Na_2Al_2O_{16}S_4$)

Smith, 1909

%	f.t.	%	f.t.
26.9	10	30.1	25
27.9	15	31.4	30
29.0	20		

Water + Potassium aluminum alum ($KAlS_2O_8$)

Lescoeur and Mathurin, 1888

t	p dissoc. (12+1)	p dissoc. (3+1)
55	72.5	-
60	98.0	-
65	128.0	-
70	164.0	-
75	210.0	-
90	-	84
100	-	127
115	-	above 320

Gerlach, 1886

%	b.t.	%	b.t.
0	100	45.62	104
14.53	100.5	47.56	104.5
23.20	101	49.39	105
29.48	101.5	50.94	105.5
34.03	102	52.49	106
37.66	102.5	53.89	106.5
40.72	103	54.66	106.7
43.40	103.5		

Poggiale, 1843

%	f.t.	%	f.t.	
2.06	0	21.07	60	
4.75	10	25.99	70	
7.18	20	31.35	80	
9.86	30	36.96	90	
12.95	40	42.71	100	(12+1)
16.73	50			

Mulder, 1866

%	f.t.	%	f.t.	
2.91	0	5.48	20.5	(12+1)
3.85	8.25	10.47	40.25	
3.94	10.5	24.53	66.25	
5.21	17.75	67.80	111.9	b.t.

Berkeley, 1904

%	f.t.	
2.91	0.40	
4.84	15.30	
7.25	28.10	
11.74	43.20	(12+1)
20.04	60.45	

Marino, 1905

%	f.t.	%	f.t.	
2.46	0	25.55	75	
3.26	5	26.63	76	
3.57	10	27.43	77	
4.28	15	28.20	78	
4.88	20	29.11	79	
5.28	25	30.08	80	
7.08	30	31.60	82	
7.71	36	33.07	84	
9.11	40	33.57	84.6	
10.23	45	34.58	85.1	(12+1)
12.24	50	35.00	85.3	
14.29	55	35.10	85.6	
16.80	60	36.40	86	
18.88	65	37.70	87	
22.09	70	39.54	88	

von Stackelberg, 1896

P	%
23°	
1	11.5
200	12.5
400	14.2

Bindel, 1890

%	d	%	d
20°			
6.33	1.0650	21.61	1.2437
11.35	.1204	30.94	.3476
13.48	.1448	39.45	.4357
16.60	.1826		

Berkeley, 1904

%	t	d
	sat.sol.	
2.91	0.40	1.0292
4.84	15.30	.0461
7.25	28.10	.0661
11.74	43.20	.1044
20.04	60.45	.1835

Kasankin, 1891

t	capillary rise
sat.sol.	
17	25.69
34	24.60

Rehberg, 1949

transparency of solutions .

Bindel, 1890

%	U	%	U
	20°		
6.33	0.943	21.61	.822
11.35	.902	30.94	.757
13.48	.885	39.45	.714
16.60	.860		

%	Q diss (by mole alum)
6.33	- 3600
11.35	- 3667
13.48	- 3721
16.60	- 3758
21.61	- 3994
30.94	- 4486
39.45	- 7754

Water + Rubidium alum ($RbAlO_8S_2$)

Berkeley, 1904

%	f.t.	d_t (sat.sol.)
0.72	0.40	1.0072
1.26	15.20	.0112
2.32	32.20	.0165
3.97	45.80	.0267
6.77	59.65	.0466
10.89	69.75	.0804

Water + Rubidium chromium sulfate ($RbCrO_8S_2$)

Rehberg, 1949

Transparency of solutions .

Water + Cesium alum ($CsAlO_8S_2$)

Berkeley and Applebey, 1911

Sat.sol. b.t. = 100.527°

Berkeley, 1904

%	f.t.	d
0.20	0.40	1.0017
0.34	15.60	1.0022
0.57	29.15	1.0010
1.06	45.25	0.9994
2.01	60.60	1.0004
3.13	75.35	1.0107
6.42	83.05	1.0250
10.11	90.85	1.0328
18.59	100.4 b.t.	1.1285

Water + Thallium alum ($TlAlO_8S_2$)

Berkeley, 1904

%	f.t.	d
3.12	0.45	1.0299
5.31	16.10	.0503
8.53	29.85	.0808
11.58	37.50	.1090
15.61	45.20	.1500
20.25	52.30	.2051
26.16	60.05	.2812

Water + Potassium ferric alum ($KFeS_2O_8$)

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987	15	1.0658
1	1.0041	16	.0702
2	.0095	17	.0746
3	.0149	18	.0790
4	.0203	19	.0834
5	.0255	20	.0880
6	.0295	21	.0928
7	.0335	22	.0976
8	.0375	23	.1024
9	.0415	24	.1072
10	.0453	25	.1122
11	.0493	26	.1179
12	.0534	27	.1236
13	.0575	28	.1293
14	.0616	29	.1350
		30	.1407

N.B. According to Gerlach, 1889, it is probably not Potassium but Ammonium ferric alum.

Gerlach, 1889

%	d	%	d
	15°		
0	0.9991	11.4152	1.1040
2.8538	1.0241	14.2690	.1330
5.7076	.0498	17.1228	.1635
8.5614	.0764	19.9760	.1957

Water + Potassium ferricyanide ($K_3FeC_6N_6$)

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
	25°		
0.1	0.727	0.7	0.679
0.2	0.695	0.8	0.685
0.3	0.682	0.9	0.694
0.4	0.678	1.0	0.705
0.5	0.676	1.2	0.727
0.6	0.676	1.4	0.750

Jones, 1904 and Jones and Bassett, 1905

M	f.t.	M	f.t.
0.05	- 0.30	0.3	- 1.57
0.1	- 0.56	0.4	- 2.05
0.2	- 1.086		

Wallace, 1909

%	f.t.	%	f.t.
24.81	4.4	37.03	37.8
26.79	10.01	43.66	100.
28.98	15.6	45.23	104.4

Berkeley and Hartley, 1915 - 16

%	P (osmotic pressure)
0°	
16.25	32.39
21.89	47.61

Friend and Smirles, 1928

%	f.t.	%	f.t.
23.22	0.1	34.27	28.3
25.49	4.7	36.65	29.8
26.96	7.8	35.15	31.5
28.68	12.8	37.22	39.9
30.35	15.7	39.12	49.0
30.96	18.7	40.41	56.25
32.08	22.1	41.10	58.0
32.80	25.0	44.70	81.0
33.66	26.3	47.60	99.0

Schiff, 1860

%	d	%	d	%	d
		13°			
0	0.9994	11	1.0595	21	1.1195
1	1.0043	12	.0653	22	.1259
2	.0097	13	.0712	23	.1324
3	.0149	14	.0771	24	.1389
4	.0202	15	.0831	25	.1455
5	.0253	16	.0891	26	.1522
6	.0309	17	.0952	27	.1589
7	.0364	18	.1014	28	.1657
8	.0420	19	.1076	29	.1725
9	.0476	20	.1139	30	.1795
10	.0530				

Sentis, 1897

mol%	t	d
0.5	12.35	1.0468
1	13.5	.0888
1	19.8	.0862
1	20.8	.0858

Chéneveau, 1907

%	d
14.8°	
0	0.9992
16.88	1.0959

Wallace, 1909

%	t	d	%	t	d
24.81	4.4	1.150	37.03	37.8	1.224
26.79	10.01	.160	43.66	100	.249
28.98	15.6	.177	45.23	104.4	.264
				b.t.	

Berkeley and Hartley, 1915 - 16

%	d
0°	
0	0.99987
16.25	1.09649
21.89	1.1319

Jones, 1904 and Jones and Bassett, 1905

M	d
0°	
0.05	1.009632
0.1	.020148
0.2	.036336
0.3	.053108
0.4	.072616

Heydweiller, 1921

N	d
18°	
0.5	1.02943
1	1.0584
2	1.1141

Friend and Smirles, 1928

%	t	d	%	t	d
		sat.sol.			
26.96	7.8	1.1567	33.66	26.3	1.1928
30.35	15.7	.1738	34.27	28.3	.1971
	16.75	.1766	35.15	31.5	.2045
32.80	25.0	.1855		33.1	.2115
32.08	22.1	.1872	40.41	56.25	.2269

Jones and Christian, 1944

N	d	N	d
25°		0°	
0.001	0.99714	0.001	0.99998
.002	.99721	.002	1.00000
.005	.99739	.005	.00021
.010	.99768	.010	.00053
.020	.99830	.020	.00111
.050	1.00009	.050	.00305
.100	.00305	.100	.00616
.29332	.01439	.29443	.01823
.300	.01478	.300	.01857
.500	.02632	.500	.03074
1.000	.05472	.50215	.03088
.00048	.05474	1.000	.06036
.99814	.11002	.00615	.06071
2.000	.11012	2.000	.11763
.99883	.16399		
3.000	.16405		
.500	.19038		
.50321	.19055		

de Lannoy, 1895

t	relative volume 4 %	8 %	12 %	20 %
0	1.00000	1.00000.	1.00000	1.00000
10	.00087	.00138	.00188	.00250
20	.00274	.00358	.00440	.00554
30	.00547	.00658	.00770	.00902
40	.00905	.01024	.01152	.01296
50	.01339	.01451	.01586	.01745
60	.01798	.01950	.02080	.02230
70	.02350	.02517	-	.02762
80	.02970	.03145	-	.03350
90	.03682	. -	-	-

Jones and Christian, 1944

N	η (water=1)	N	η (water=1)
	25°		
0.001	1.00047	0.500	1.03164
0.002	.00065	1.000	.06678
0.005	.00121	1.00048	.06682
0.010	.00179	1.99814	.17067
0.020	.00280	2.0000	.1709
0.050	.00503	2.99883	.33648
0.100	.00828	3.000	.3367
0.29330	.01939	3.500	.4489
0.300	.01978	3.50321	.44964
	0°		
0.001	1.00038	0.29443	0.99224
0.002	.00045	0.300	.99210
0.005	.00059	0.500	.98818
0.010	.00069	0.50215	.98820
0.020	.00055	1.000	.98879
0.050	0.99985	1.00615	.98891
0.100	.99820	2.000	1.03289

Sentis, 1897

mol %	t	σ
0	25.1	72.5
0	13.5	74.0
0.5	12.35	75.3
1	20.8	74.9
1	19.8	75.3
1	13.5	75.9

Gellings, 1956

M	d	λ	M	d	λ
		25°			
0.030658	-	129.43	1.4407	1.07947	90.22
.047887	0.99999	123.94	.5010	.08278	89.72
.092663	-	116.15	.7587	-	87.70
.191260	1.00842	108.61	.9988	1.11022	85.95
.312370	.01560	104.13	2.2012	.12121	84.40
.368770	.01893	102.64	.3517	.12932	83.24
.562050	.02992	99.03	.5809	-	81.44
.825700	.04497	95.78	.8676	-	79.18
.976830	.05351	94.19	.9438	1.16110	78.65
1.252400	-	91.73	.9567	.16184	78.56

Chéneveau, 1907

%	n C	D	Tl	F
		14.8°		
0	-	1.33338	-	-
16.88	1.36012	.36238	1.36471	1.36824

Jones, 1904 and Jones and Bassett, 1905

M	
0°	
0.05	172.4
0.1	162.8
0.2	154.3
0.3	149.2
0.4	146.1

Heydweiller, 1921

N	λ
18°	
0.5	86.5
1	82
2	75

Gray and Birse, 1914

%	χ	%	χ
	at room temperature		
0	-0.72	12.37	+0.162
2.976	.506	15.55	.391
4.594	.383	17.32	.513
6.005	.286	20.35	.731
7.416	.178	23.53	.993
8.27	.123	26.22	1.188
11.08	+0.0675	29.13	.394

Water + Yttrium chloride (YCl_3)

Crew, Steinert and Hopkins, 1925

%	f.t.	%	f.t.
42.4	0	43.4	45
43.1	16	43.6	60
43.0	25	43.9	80

Williams, Fogg and James, 1925

%	f.t.
43.85	10
44.08	20
44.33	30
44.70	40
45.07	50

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.789	0.9	1.161
.2	.810	1.0	.223
.3	.847	.2	.354
.4	.892	.4	.491
.5	.939	.6	.631
.6	.989	.8	.780
.7	1.042	2.0	.940
.8	1.100		

Water + Yttrium bromide (YBr_3)

Crew, Steinert and Hopkins, 1925

%	f.t.
39.0	0
43.5	30
49.2	50
52.7	75
56.4	95

Water + Yttrium nitrate (YN_3O_9)

Crew, Steinert and Hopkins, 1925

%	f.t.
48.3	0
57.4	22.5
60.8	35
66.7	60.2
67.9	66.5

Water + Ytterbium sulfate ($Yb_2O_{12}S_3$)

Cleve, 1902

%	f.t.	%	f.t.
4.46	70	9.42	60
5.53	80	9.95	55
6.48	90	16.04	35
6.73	100	25.71	15.5
		30.65	0

Spedding and Jaffé, 1954

N	λ	N	λ
	25°		
0.000	145.2	0.01000	50.53
.0001000	131.9	.03335	37.27
.0003335	109.9	.05502	33.62
.0006252	98.79	.06669	31.35
.00100	89.67	.10000	28.40
.003335	66.63	.33350	21.51
.005002	60.05	3.33450	9.884
.006669	55.50		

Water + Samarium chloride ($SmCl_3$)

Williams, Fogg and James, 1925

%	f.t.
48.02	10
48.29	20
48.60	30
49.20	40
49.98	50

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.789	0.9	1.128
.2	.809	1.0	.186
.3	.841	.2	.302
.4	.879	.4	.427
.5	.921	.6	.554
.6	.964	.8	.686
.7	1.019	2.0	.824
.8	1.074		

Water + Samarium nitrate (SmN_3O_9)

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.0411	70.9
1.0	.0812	60.8
2.0	.1587	46.45
3.0	.2348	35.25

Tollert, 1939

N	d	η
	20°	
1.080	1.09734	1204.1
0.217	.01846	1043.7
0.0217	.00035	1009.9
0.0043	0.99873	1006.0

Jauch, 1921

N	U
18°	
0.5	0.9435
1	.8960
2	.8166
3	.7510

Water + Samarium selenate ($Sm_2O_{12}Se_3$)

Friend, 1941

%	f.t.	%	f.t.
	(8+1)		
33.60?	16.6	22.15	50.0
34.30	17.0	21.13	55.0
31.01	25.0	20.38	64.6
27.52	35.2	17.65	65.0
24.85	45.0	14.03	80.0

Water + Gallium perchlorate ($GaCl_3O_{12}$)

Patterson, Tyree, jr. and Knox, 1955

Isopiestic solutions.

m_1	m_2	m_1	m_2
	25°		
0.0549	0.1020	0.6264	1.742
0.0818	0.1542	0.6514	1.843
0.1038	0.1969	0.7248	1.860
0.1539	0.2982	0.7649	2.147
0.1955	0.3886	0.9248	3.100
0.2341	0.4744	0.9408	3.179
0.2559	0.5361	1.065	3.862
0.2930	0.6355	1.168	4.460
0.3172	0.6971	1.184	4.562
0.3319	0.7384	1.219	4.770
0.3557	0.8063	1.511	2.790
0.4181	0.9940	1.825	3.472
0.4733	1.173	2.002	3.857
0.5009	1.269		

m_1 = m of $GaCl_3O_{12}$ m_2 = m of KCl

Water + Indium chloride ($InCl_3$)

Ensslin, Ziemeck and Schaepdryver, 1949

%	f.t.	d
	(4+1)	sat.sol.
62.50	2	1.8467
65.44	14.5	1.9386
66.11	22	1.9702
67.92	24	2.0157
69.62	26	2.0298
71.12	31	2.0791
73.07	35	2.1129
	(3+1)	
69.85	30.5	2.0557
71.78	41	2.0797
71.64	41	2.0833
73.31	51	2.1668
75.34	60	2.1961
77.71	70	2.2239
78.14	75	2.2304
78.46	75	2.2287
78.37	75	2.2288
	(5+2)	sat.sol.
78.87	80	2.2466
78.79	81	2.2509
79.23	85	2.2563
79.50	90	2.2522
80.27	94	2.2727
80.74	98	2.2793
	(2+1)	
81.33	105	2.2894
81.35	110	2.2901
81.68	118	2.3028

Water + Indium bromide ($InBr_3$)

Ensslin, Ziemeck and Schaepdryver, 1949

%	f.t.	%	f.t.
	(5+1)		
71.26	0	76.71	18
71.66	2	79.94	19
72.53	4	80.55	20
74.51	14	84.64	22
75.62	16		
	(2+1)		
85.26	22	85.19	29
85.13	23	85.35	29.5
85.10	25.5		
	anh.		
85.19	34.5	86.49	70
85.53	40	87.48	100
85.75	50	87.91	105
86.09	60		

%	f.t.	d
	(5+1)	sat.sol.
71.26	0	2.3773
71.66	2	.3838
72.53	4	.3935
74.51	14	.4593
75.62	16	.4727
76.71	18	.5505
79.94	19	.6100
80.55	20	.6668
84.64	22	.8313
	(2+1)	
85.26	22	2.8291
85.13	23	.8299
85.10	25.5	.8350
85.19	29	.8512
85.35	29.5	.8519
	anh.	
85.19	34.5	1.8675
85.53	40	.8683
85.75	50	.8936
86.09	60	.8974
86.49	70	.8975
87.48	100	.8880
87.91	105	.8939

Heydweiller, 1914

N	d	N	d
	18°		
0	0.99862	1.0038	1.09672
0.1003	1.00854	1.0038	.19301
.2000	.01851	3.7875	.36534
.4992	.04861		

N	n Hα	n Hβ	n Hγ
		18°	
0.	1.33140	1.33739	1.34054
0.4992	.33909	.34546	.34890
1.0038	.34670	.35339	.35705
1.9986	.36156	.36894	.37304
3.7875	.38820	.39684	.40181

N	κ	N	κ
	18°		
0.1003	54.0	1.0038	287.8
.2000	93.5	1.9986	396.4
.4992	184.8	3.7875	418.5

N	α_D (equivalent)
18°	
1.0038	9.346
1.9986	9.286
3.7875	8.934

Jauch, 1921

N	U
18°	
0.5	0.9418
1	.8922
2	.8020
3	.7285
4	.6652

Water + Indium iodide (InI_3)

Ensslin, Ziemeck and Schaepdryver, 1949

%	f.t.	d	%	f.t.	d
		sat.sol.			sat.sol.
92.31	1	3.4432	94.23	40	3.4749
92.35	4.5	.4443	94.24	51	.4921
92.40	7	.4462	94.48	61	.5003
92.91	22	.4560	94.81	69	.5212
93.51	30	.4627	95.29	70	.5238

Water + Indium sulfate ($In_2S_3O_{12}$)

Lietzke and Stoughton, 1956

m	t	e	m	t	e
0.009888	13.2	1.1048	0.5527	14.6	1.0830
	20.4	.0977		33.2	.0661
	27.5	.0905		42.45	.0577
	35.1	.0837		50.5	.0502
0.04933	14.5	.0992		58.3	.0434
	23.9	.0901		68.9	.0342
	28.3	.0862		74.6	.0297
	35.6	.0795	0.6806	13.6	.0791
0.09799	14.3	.0960		22.7	.0706
	24.4	.0870		29.8	.0645
	40.3	.0730		30.4	.0640
	50.2	.0647		39.5	.0558
0.1957	24.9	.0841		50.6	.0468
	36.2	.0735		62.3	.0370
	44.8	.0661		71.5	.0299
	50.4	.0616	0.9732	24.9	.0599
	60.3	.0533		48.2	.0422
0.1990	15.0	.0923		63.4	.0310
	20.0	.0880	1.2909	14.0	.0588
	21.1	.0836		19.3	.0531
0.3959	13.3	.0915		29.5	.0449
	20.2	.0838		40.8	.0368
	30.0	.0747		50.2	.0306
	35.3	.0701		60.3	.0245
	41.8	.0644		70.5	.0176
	49.9	.0576	2.2822	12.5	.0296
	57.3	.0518		21.8	.0225
	65.4	.0450		23.1	.0215
	70.5	.0409		29.0	.0175
				39.4	.0105
				40.2	.0102
				49.8	.0053
				70.5	.9951

m	κ
25°	
0.01	24.2
.03	52.7
.10	128
.30	267
1.00	347

Water + Indium sesquioxide (In_2O_3)

Roy and Shafer, 1954 (fig.)

t	PKg dissoc.	t	PKg dissoc.
(3+1)-(1+1)		(1+1)-anh.	
250	80	410	80
260	190	425	110
265	380	430	500
270	1000	440	700
		449	900

Water + Scandium chloride ($ScCl_3$)

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.797	0.8	1.156
.2	.827	0.9	.222
.3	.868	1.0	.291
.4	.917	.2	.430
.5	.969	.4	.572
.6	1.027	.6	.718
.7	1.090	.8	.869

Water + Europium chloride ($EuCl_3$)

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.794	0.9	1.137
.2	.812	1.0	.193
.3	.842	.2	.310
.4	.882	.4	.438
.5	.926	.6	.570
.6	.971	.8	.707
.7	1.027	2.0	.853
.8	1.137		

Water + Neodymium chloride ($NdCl_3$)

Williams, Fogg and James, 1925

%	f.t.
49.16	10
49.49	20
49.91	30
50.41	40
51.11	50

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.783	0.9	1.110
.2	.801	1.0	.165
.3	.832	.2	.283
.4	.871	.4	.404
.5	.913	.6	.527
.6	.954	.8	.656
.7	1.006	2.0	.789
.8	1.056		

Selwood, 1930

%	t	d	n^{20}_{6362}
13.46	26.6	1.1357	1.35898
20.40	26.9	.2198	.37570
25.37	27.0	.2890	.38922
30.53	27.6	.3664	.40356
35.01	27.8	.4393	.41724
38.18	24.6	.5001	.42873
44.84	25.2	.6283	.45228
47.55	25.0	.6936	.46360

Water + Neodymium nitrate (NdN_3O_9)

Friend, 1935

%	f.t.	%	f.t.	
		(6+1)		
55.97	0	59.17	27.2	II
57.37	13.2 I	59.18	29.4	
58.03	18.2	60.95	37.2	
59.59	23.0	61.91	42.4	
60.46	25.0	64.86	50.0	
60.69	26.2	67.00	57.2	
		71.13	66.2	
		75.34	67.5	

Selwood, 1930

%	t	d	n_{6563}
11.92	24.6	1.1054	1.34858
19.14	24.8	.1792	.35966
26.50	24.3	.2652	.37268
34.49	24.2	.3698	.38800
38.56	25.2	.4290	.39637
42.95	25.2	.4930	.40530
48.32	25.7	.5867	.41812
52.76	25.8	.6620	.42815
58.02	26.0	.7604	.44115
60.10	25.8	.7986	.44604

Tollert, 1939

N	d	η
	20°	
1.00	1.09515	1198.8
0.10	.00820	1025.3
.02	.00031	1009.7
.004	0.99873	1006.1

Water + Neodymium selenate ($Nd_2O_{12}Se_3$)

Friend, 1931

%	f.t.	%	f.t.	
31.10	0.0	28.34	30.8	
30.84	0.0	27.90	43.4	
30.79	8.6	29.44	51.5	
30.30	12.6	29.30	54.0	
29.25	13.0	30.00	55.4	
28.27	13.4	30.67	58.4	(5+1)
29.81	15.0	26.14	61.6	(6+1)
30.85	17.0	19.44	63.2	
29.16	17.6	17.13	64.6	(8+1)
27.91	18.4	16.8	65.2	(12+1)
28.92	23.6	8.33	77.0	
29.38	26.2	6.98	79.4	(18+1)
29.46	30.0	3.19	82.4	(22+1)

Water + Praseodymium chloride ($PrCl_3$)

Robinson and Stokes, 1949

m	osmotic coefficient	m	osmotic coefficient
		25°	
0.1	0.784	0.9	1.100
0.2	0.801	1.0	1.154
0.3	0.830	1.2	1.271
0.4	0.866	1.4	1.388
0.5	0.905	1.6	1.507
0.6	0.945	1.8	1.631
0.7	0.996	2.0	1.759
0.8	1.046		

Water + Praseodymium nitrate (PrN_3O_9)

Friend, 1935

%	f.t.	%	f.t.
		(6+1)	
59.32	15.8	65.00	43.0
60.18	22.0	75.15	56.0
61.94	30.4		

Tollert, 1939

N	d	η
	20°	
1.00	1.09423	1196.9
0.10	.00814	1024.8
.02	.00029	1010.4
.004	0.99865	1006.4

Water + Praseodymium sulfate ($Pr_2S_3O_{12}$)

Spedding and Jaffé, 1954

9.81 % f.t. = 25°

N	λ	N	λ
	25°		
0.01148	44.50	0.1148	24.18
0.02869	35.05	0.5739	15.84
0.05739	29.12	1.1477	12.81
0.08608	26.05		

Water + Praseodymium selenate ($Pr_2O_{12}Se_3$)

Friend, 1932

%	f.t.	%	f.t.
(x+1)		(12+1)	
26.59	0.5	23.33	59.5
23.88	17.4	14.00	63.6
23.79	17.8	9.96	67.0
24.29	24.6	6.89	75.0
24.50	30.0	5.64	81.0
23.67	40.2	2.99	92.0
24.00	48.6		
23.89	52.0		
23.99	55.0		

Water + Praseodymium ammonium nitrate ($PrH_8N_7O_{15}$)

Jones and Caldwell, 1901

c	f.t.
56.840	-16.5
28.420	-7.5
14.210	-3.637
5.684	-1.779
2.842	-0.887

c	molar conduct.	c	molar conduct.
	25°		
56.840	134.1	0.02842	555.8
28.420	248.0	.01421	580.3
14.210	334.9	.05684	613.2
5.684	404.7	.02842	628.0
2.842	451.1	.01421	642.8
1.421	491.1	.0568	654.3
0.5684	531.3		

Water + Erbium sulfate ($Er_2O_{12}S_3$)

Spedding and Jaffé, 1954

N	λ	N	λ
		25°	
0.0000	145.9	0.01454	44.27
.0002907	110.3	.02907	37.18
.0005814	97.96	.05814	31.16
.0008721	90.01	.08721	28.17
.001454	79.60	.1644	24.02
.002907	66.20	.2907	20.96
.005814	55.90	1.4536	13.66
.008721	50.42		
		13.19%	f.t.=25°

Water + Lanthanum chloride ($LaCl_3$)

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
		25°	
0.1	0.788	0.9	1.102
.2	.800	1.0	.154
.3	.833	.2	.266
.4	.871	.4	.384
.5	.912	.6	.502
.6	.955	.8	.623
.7	.998	2.0	.748
.8	1.052		

Jones and Bickford, 1934

M	d	molar. conduct.	M	d	molar. conduct.
	25°			0°	
0.00025	0.99713	414.61	0.00025	0.99993	216.47
.00036	.99715	410.11	.00036	.99996	214.36
.0005	.99718	405.64	.0005	.99999	212.15
.00075	.99724	398.83	.00075	1.00005	208.93
.001	.99727	393.48	.001	.00009	206.29
.0025	.99762	372.84	.0025	.00044	196.21
.005	.99820	354.33	.005	.00104	187.15
.01	.99932	333.74	.1	.00225	176.98
.025	1.00274	305.90	.025	.00575	163.23
.05	.00839	284.86	.05	.01152	152.85
.10	.01960	263.68	.1	.02297	142.62
.25	.05292	231.99	.25	.05709	127.04
.5	.10756	200.03	.5	.11273	111.12
1.0	.21455	153.44	1.0	.22092	86.31

Jones and Stauffer, 1940

M	d	η(water=1)	M	d	η(water=1)
	25°			0°	
0.00025	0.99713	1.00067	0.00025	-	1.00051
.0005	.99718	.00091	.0005	-	.00072
.001	.99729	.00152	.001	1.00011	.00129
.0025	.99763	.00291	.0025	.00047	.00266
.005	.99820	.00499	.005	.00107	.00444
.01	.99933	.00879	.01	.00221	.00731
.025	1.00275	.01891	.025	.00570	.01614
.05	.00839	.03517	.05	.01150	.03003
.1	.01965	.06632	.1	.02302	.05621
.25	.05299	.16314	.250395	.05717	.13872
.502687	.10812	.35132	.250462	-	.13876
.990638	.21256	.83815	.505272	.11443	.30640
			.997033	.22039	.76661

Guillaume, 1946

%	d	t	(α)* magn.10^6 5780 Å	n
		20°		
9.10	1.0964	20	-	1.3544
29.1	1.3728	16	4.020	1.4092

*in radians, gauss, centim.

Mason, Gray and Ernst, 1936

M	Verdet's constant.10^3 10°	20°	25°	30°	40°	50°
			5893 Å			
0.00	13.19	13.12	13.07	13.10	13.10	13.10
.10	13.46	13.46	13.47	13.53	13.53	13.53
.40	14.40	14.40	14.37	14.40	14.33	14.33
.80	15.67	15.60	15.55	15.40	15.33	15.33
.9967	16.07	16.07	16.08	16.14	16.14	16.14
1.40	17.21	17.21	17.22	17.21	17.14	17.14
1.7509	-	18.08	18.21	18.15	18.01	17.94
2.20	19.08	19.15	19.14	19.10	19.10	19.10
2.50	19.89	19.86	19.83	19.69	19.69	19.69
2.80	20.56	20.56	20.35	20.30	20.29	20.15
3.10	-	20.96	20.97	20.96	20.96	20.96
3.40	21.56	21.56	21.52	21.43	21.29	21.23
3.9340	21.96	21.96	22.03	21.96	21.76	21.69
			5460.7 Å			
0.00	15.40	15.40	15.40	15.40	15.40	15.40
.10	15.73	15.73	15.75	15.80	15.80	15.80
.40	17.07	16.94	16.85	15.87	16.87	16.87
.80	18.35	18.35	18.17	18.40	18.41	18.41
.9967	19.15	19.08	19.03	19.02	19.02	19.02
1.40	20.42	20.35	20.35	20.29	20.22	20.22
1.7509	-	21.43	21.49	21.49	21.43	21.43
2.20	22.43	22.50	22.50	22.43	22.43	22.43
2.50	23.37	23.34	23.28	23.30	23.30	23.30
2.80	24.10	24.07	24.04	24.00	23.95	23.90
3.10	-	24.90	24.80	24.64	24.57	24.31
3.40	25.44	25.44	25.43	25.31	25.24	25.11
3.9640	26.18	25.91	25.88	25.85	25.80	25.71

Water + Lanthanum nitrate (LaN_3O_9)

Di Capua, 1929

60.13% f.t.= 20°

Friend, 1935

%	f.t.	%	f.t.
(6+1) I		(6+1) II	
50.03	0	56.27	14.4
54.16	18.4	56.94	15.2
55.03	21.2	56.85	15.8
55.80	25.4	56.74	16.0
59.12	35.4	56.70	23.2
63.84	42.4	60.08	29.6
65.13	44.2	61.34	32.2
		62.71	40.0
		64.55	46.4
		65.17	49.4
		68.30	56.0
		75.04	65.4

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.04798	63.35
1	.0951	54.04
2	.1866	39.13
3	.2762	28.45
4	.3642	19.92

Tollert, 1939

N	d	η	N	d	η
			20°		
3.232	1.27680	1843.6	0.1014	1.00839	1025.9
1.268	.121990	1265.2	.0101	0.99908	1007.9
0.507	.04830	1098.4	.0051	0.998585	1006.3

Jauch, 1921

N	U
0.5	0.9395
1	.8880
2	.8036
3	.7321
4	.6688

Water + Lanthanum selenate ($La_2Se_3O_{12}$)

Friend, 1932

%	f.t.	%	f.t.
(x+1)		(12+1)	
33.55	0.0	34.00	36.4
32.87	0.6	27.54	46.2
30.90	9.6	23.37	51.4
30.84	15.0	15.92	59.4
31.31	21.8	8.51	69.4
30.54	25.4	5.02	78.2
31.29	33.6	3.78	81.6
31.15	40.6	1.93	92.4
		1.97	93.4

Water + Titanium trichloride ($TiCl_3$)

Heydweiller, 1915

M	ϰ initial	ϰ after a long time
	18°	
1.228	1189	2200
0.609	969	1755
0.297	582	1050

Water + Cerous chloride ($CeCl_3$)

Tammann, 1885

%	p	%	p
	100°		
6.07	749.7	26.90	676.2
16.00	723.7	35.11	605.6
18.74	712.9	48.72	455.0

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.782	0.9	1.107
.2	.805	1.0	.158
.3	.835	.2	.264
.4	.872	.4	.387
.5	.914	.6	.504
.6	.955	.8	.638
.7	1.007	2.0	.777
.8	1.057		

Gibson, 1935

%	d	π (1-1000bars)
	25°	
0.00	0.9970	39.35
8.69	1.0832	35.03
16.97	.1746	30.93
23.36	.2561	28.00
34.60	.4300	22.75
44.09	.6041°	18.80

Slack, Reeves and Pegles, 1934

c	$d_{20°}$	n_{D} 26°-30°	c	$d_{20°}$	n_{D} 26-30°
0	1.000	1.3324	35.92	1.3417	1.4015
6.41	.0619	.3449	46.70	.4393	.4211
14.09	.1335	.3594	49.35	.4575	.4257
23.95	.2270	.3792	64.15	.5872	.4476
26.31	.2449	.3920			

Slack, Reeves and Pegles, 1934

c	Verdet's constant.10^3 (in minutes) 15°	20°	30°	40°	45°
			5460.7 Å		
64.15	-31.35	-30.40	-28.80	-27.60	-26.95
49.35	-18.60	-17.70	-16.90	-16.15	-15.80
46.70	-16.90	-16.30	-15.25	-14.20	-13.70
35.92	-8.85	-8.35	-7.42	-6.43	-5.95
26.31	-1.75	-1.45	-0.85	-0.35	-0.10
23.31	-0.35	0.00	+0.60	+1.25	+1.60
23.95	+6.42	+6.60	+6.90	+7.25	+7.40
14.09	+11.45	+11.50	+11.70	+11.80	+11.90
6.41	+15.65	+15.60	+15.46	+15.35	+15.30
			4481 Å		
64.15	-42.75	-41.55	-39.20	-36.80	-35.60
49.35	-28.00	-26.75	-25.10	-23.40	-22.75
26.31	-5.85	-5.40	-4.45	-3.45	-3.00
			D		
64.15	-25.6	-24.70	-23.30	-22.10	-21.50
49.35	-14.85	-14.40	-13.60	-12.70	-12.28
26.31	-1.10	-0.80	-0.35	+0.20	+0.45
0.00	+13.20	+13.19	+13.20	+13.18	+13.17

Tammann and Röhmann, 1929

P Kg	t	m 0.5 N	1.0 N	5.0 N
800	20	4.95	4.73	-0.07
	40	3.12	2.99	-
1000	20	8.38	8.23	-2.33
	40	5.19	4.91	-
1500	20	10.30	10.20	-4.47
	40	6.30	6.12	-
2000	20	11.20	11.10	-6.95
	40	6.79	6.55	-
2500	20	11.60	11.20	-9.75
	40	6.82	6.54	-
3000	20	11.20	10.80	-12.80
	40	6.59	6.24	-

$x = 100\ (m_P - m_{P=1}) / m_{P=1}$ where m= molecular conductivity

Water + Cerous nitrate (CeN_3O_9)

Tollert, 1939

N	d	η
	20°	
2.000	1.16988	1403.6
1.000	1.08515	1176.5
0.500	1.04216	1087.0
0.100	1.00706	1022.3
0.010	0.99903	1007.8
0.002	0.99828	1005.7

Water + Cerous sulfate ($Ce_2O_{12}S_3$)

Muthmann and Rölig, 1898

%	f.t.	%	f.t.
	(5+1)		
0.77	100	5.26	50
1.68	80	7.58	40
3.33	60		
	(8+1)		
16.04	0	11.09	50
14.76	18	8.59	60
13.89	30	4.07	70
	(12+1)		
17.62	0	13.95	25
15.57	18		

Voogd, 1933

% (8+1)	(9+1)	f.t.
15.85	17.60	0
8.63	8.95	20.0
7.59	7.72	25.0
-	6.76	30.0
6.63	-	30.4
5.99	5.95	35.0
5.40	5.34	40.0
4.90	4.79	45.0

Brauner, 1888

%	d	%	d
anh.	15°	(8+1)	
0	0.99913	3.07	1.03008
3.07	1.03005	5.03	.05996
5.75	.05812	7.69	.07910
7.80	.08000	7.82	.08031
8.75	.09085	9.51	.09928
9.45	.09939	9.52	.09959
11.23	.11917	10.34	.10987
12.70	.13665	10.90	.11530
13.54	.14623	12.66	.13618
17.48	.19640		
24.05	.28778		

Tollert, 1939

N	d	η
	20°	
0.9045	1.06210	1176.0
.1758	.01101	1035.8
.0351	.00089	1011.0
.0070	0.99886	1006.5

Gibson, 1934

mol%	d	π (1-1000 bars)
	25°	
0	0.9970	-
8	1.0791	33.96

Water + Cerous ammonium nitrate ($CeH_8N_7O_{15}$)

Wolff, 1905

%	f.t.
(4+1)	
70.2	8.75
74.8	25.00
80.4	45.00
87.2	60.00
89.1	65.06

Water + Ferric chloride ($FeCl_3$)

Heterogeneous equilibria .

Roozeboom, 1892

mol %	f.t.	mol %	f.t.
0	0	1.86	-27.5
0.99	-10	2.31	-40
1.61	-20.5	2.67	-55
	(12+1)		
2.67	-55	7.69	37
2.72	-41	8.50	36
2.89	-27	9.47	33
3.97	0	10.07	30
4.34	+10	10.83	27
4.85	20	11.37	20
5.60	30	11.66	10
6.35	35	12.05	8
7.35	36		
	(7+1)		
10.11	+20	12.40	32.5
10.83	27.4	13.13	30
11.93	32.0	13.45	25
	(5+1)		
11.39	+12	14.89	50
12.24	20	16.08	55
12.92	27	16.67	56
13.13	30	16.89	55
13.52	35		
	(4+1)		
16.64	+50	20.00	73.5
16.89	55	20.72	72.5
17.15	60	21.82	70
17.71	69	22.60	66
18.94	72.5		
	anh.		
22.60	+66	22.60	80
22.73	70	22.60	100
22.43	75		

f.t. or tr.t.	solid phase
37	(12+1)
32.5	(7+1)
56	(5+1)
73.5	(4+1)
55	E
27.4	(12+1)-(7+1)
30	(7+1)-(5+1)
15	(12+1)-(5+1)
55	(5+1)-(4+1)
66	(4+1)- anh.

p	solid phase
6.0	(12+1)
1.4	(12+1)
1.8	(12+1)-(7+1)
2.3	(7+1)
1.3	(5+1)

Mohr, 1897

mol%	f.t.	
9.89	25	(6+1)
11.78	35	

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.064	-0.387	1.287	-12.940
.103	-0.607	1.544	-17.650
.129	-0.758	2.058	-30.500
.257	-1.578	2.573	-51.000
.515	-3.688		

Jones and Stine, 1909

M	%	f.t.
0.2	3.159	- 1.255
0.4	6.160	- 2.715
0.6	9.015	- 4.530
0.8	11.750	- 6.667
1.0	14.36	- 9.152
1.2	16.85	-12.110
1.36	18.80	-14.592

Malquori, 1928

49.76% f.t. = 25°

Linke, 1956

%	f.t.		%	f.t.	
22.7	-20.5		35.74	-9.8	
24.5	-23.9		37.34	-6.4	
24.7	-24.8		39.60	-3.2	
26.0	-27.7		42.00	-0.4	
27.9	-32.1		42.40	0.0	tr.t.
29.8	-38.3		38.35	-20.0	(6+1)
28.7	-35.0	E	42.40	0.0	
29.28	-31.6	(10+1)	49.38	25.0	
32.0	-20.0				

Properties of phases

Schult, 1869

%	d 4.8°	d 9.7°	d 14.6°	d 19.7°
0	0.9998	0.9992	0.9983	0.9971
2.79	-	-	1.0220	-
4.65	-	-	.0372	-
10.45	1.0939	1.0930	1.0918	1.0901
16.79	.1534	.1521	.1507	.1491
22.54	.2140	.2129	.2107	.2090
24.60	.2351	.2334	.2318	.2298
33.25	.3381	.3359	.3339	.3317
36.95	-	.3847	.3824	.3800
41.00	.4413	.4387	.4361	.4335
49.61	.5609	.5585	.5540	.5497

60% $d^{25} = 1.670$

Franz, 1872

%	d	%	d
	17.5°		
0	0.999	40	1.413
10	1.086	50	.545
20	.178	60	.668
30	.290		

Hager, 1876

%	d	%	d
	17.5°		
1	1.007	31	1.302
2	.015	32	.314
3	.024	33	.326
4	.032	34	.338
5	.041	35	.350
6	.050	36	.362
7	.059	37	.374
8	.068	38	.388
9	.077	39	.401
10	.086	40	.413
11	.094	41	.426
12	.103	42	.439
13	.112	43	.452
14	.122	44	.467
15	.130	45	.479
16	.139	46	.492
17	.149	47	.505
18	.158	48	.518
19	.168	49	.531
20	.178	50	.545
21	.189	51	.558
22	.200	52	.571
23	.210	53	.585
24	.221	54	.598
25	.232	55	.610
26	.243	56	.622
27	.254	57	.634
28	.266	58	.646
29	.278	59	.657
30	.290	60	.668

Jones, 1904 and Jones and Getman, 1904

M	d	M	d	M	d
0.064	1.011104	0.257	1.036560	1.544	1.196836
0.103	.016612	0.515	.070524	2.058	.258120
0.129	.020220	1.287	.167696	2.573	.318164

Blümcke, 1884

%	t	d	%	t	d
0	16	0.99916	28.88	16	0.27820
20.00	18	.17820	43.56	15.5	.46100

Jones and Stine, 1909

M	%	d	M	%	d
		0°			
0.0	0	0.999868	0.8	11.75	1.104920
0.2	3.159	1.027168	1.0	14.36	.129600
0.4	6.160	.053520	1.2	16.85	.155468
0.6	9.015	.079824	1.361	18.80	.174760

Cabrera and Moles, 1912

%	d	%	d	%	d
			19°		
0	0.9984	0.8247	1.0056	4.8260	1.0395
0.2105	1.0006	1.2351	.0084	9.0630	.0762
0.4162	.0019	2.3864	.0183	14.8630	.1284

Moles, Marquina and Santos, 1913

%	d			
	0°	18°	25°	35°
0	0.9998	0.9986	0.9970	0.9940
0.4088	1.0035	1.0022	1.0003	0.9976
0.8176	.0072	.0054	.0039	1.0010
1.6975	.0139	.0121	.0106	.0075
3.178	.0276	.0255	.0239	.0208
8.737	.0774	.0740	.0718	.0684
10.934	.0984	.0946	.0923	.0889
16.199	.1498	.1451	.1428	.1390
20.196	.1904	.1853	.1828	.1789
25.729	.2465	.2411	.2381	.2339
35.275	.3654	.3573	.3541	.3488
43.579	.4705	.4654	.4611	.4548
satd	-	-	.7934	.7849

Heydweiller, 1921

c	d	c	d
	18°		
0.5	1.02263	4	1.1690
1	.04456	5	.2090
2	.0870	6	.2486
3	.1284	8	.3270

Manchot, Jahrstorfer and Zepter, 1924

%	d
25°	
13.433	1.1030
21.558	1.1638

Roth and Flügge, unpublished

%	d	%	d
	20°		
5.15	1.0424	39.16	1.4027
14.40	.1246	46.04	.4964
20.36	.1831	47.91	.5244
33.90	.3370		

Kangro and Flügge, 1936

%	d	%	d
	20°		
5.2	1.0424	34.1	1.3389
14.4	.1246	39.2	.4027
19.2	.1710	45.8	.4920
21.5	.1951	46.3	.5008
33.8	.3350	47.8	.5228

Moles, Marquina and Santos, 1913

%	0°	18°	25°	35°
0	1797	1055	894.9	724.8
0.4088	1829.9	1075	912.5	737.1
0.8176	1856.4	1095	929.3	749.5
1.6975	1928.5	1129	955.5	774.2
3.178	2052.1	1208	1029	824.4
8.737	2648.9	1538	1304	1035.8
10.934	2955	1718	1437	1141.9
16.199	3937	2227	1855	1448.8
20.196	5063	2758	2298	1765.3
25.729	7117	3744	3051	2323.8
35.275	1567.2	7308	5713	4146.0
43.579	3520.8	13770	10515	7202.0
satd	-	-	31648	19420.0

Smits, Lande and Bouman, 1921

%	min. of flow	%	min. of flow
	40°		
0	1.70	60.18	17.50
33.50	6.15	62.44	16.90
39.58	8.70	65.09	16.95
46.13	12.50	66.84	17.50
50.19	16.45	68.45	18.60
52.57	18.05	71.06	21.73
55.39	19.32	72.88	24.63
57.45	18.66	73.83	28.17
58.69	17.92	74.12	31.80

Jones, 1904 and Jones and Getman, 1904

M	n_D	M	n_D
0.064	1.32830	1.287	1.37837
0.103	.32997	1.544	.38902
0.129	.33256	2.058	.40963
0.257	.33684	2.573	.42360
0.515	.34710		

Jones and Stine, 1909

M	%	κ	M	%	κ
		0°			
0.2	3.159	238.0	1.0	14.36	531.47
0.4	6.160	378.6	1.2	16.85	531.45
0.6	9.015	464.5	1.36	18.80	517.77
0.8	11.750	514.7			

Heydweiller, 1921

N	λ	N	λ
	18°		
0.5	66.5	4	20.50
1	52.9	5	15.90
2	37.57	6	12.40
3	28.1	8	6.20

Cabrera and Moles, 1912

%	χ	%	χ
	19°		
0.2105	-0.5545	2.3864	+1.3025
0.4162	-0.3865	4.8260	3.4670
0.8247	-0.0541	9.0630	7.2547
1.2351	+0.2924	14.8630	12.4110

Thermal constants .

Roth and Flügge, unpublished

%	U	%	U
	20°		
5.15	0.935	39.16	0.631
14.40	0.833	46.04	0.588
20.36	0.773	47.91	0.579
33.90	0.663		

Blümcke, 1884

%	t	U
20.00	18	0.813
28.88	16	0.745
43.56	15.5	0.670

Kangro and Flügge, 1936

%	U	%	U	%	U
5.2	0.935	33.8	0.663	45.8	0.590
14.4	0.833	34.1	0.662	46.3	0.586
19.2	0.784	39.2	0.631	47.8	0.581
21.5	0.762				

% initial	% final	Q dil. (by mole salt)
5.15	5.67	30.60
14.40	14.91	26.71
19.20	19.51	25.00
21.52	21.92	23.94
33.75	34.00	19.36
34.05	34.46	19.31
39.36	39.50	17.69
45.78	45.93	14.90
46.30	46.69	14.31

Moles, Marquina and Santos, 1913

%	0°	18°	25°	35°
0.4088	46.89	77.2	9.05	110.9
0.8176	82.65	136.0	158.50	193.2
1.6975	142.90	230.8	267.00	321.9
3.178	241.60	383.4	441.50	527.6
8.737	467.40	722.5	823.70	973.0
10.934	517.10	797.5	908.20	1074.0
16.199	555.90	862.5	981.90	1172.0
20.196	532.90	839.7	965.00	1149.8
25.729	451.70	733.7	852.60	1033.0
35.275	248.60	452.5	543.70	681.3
43.579	-	240.4	299.00	-
satd	17.30	-	83.90	129.0

Beetz, 1879

d^{20}	heat conductivity	t
1.126	421	8-14
-	658	28-36
1.244	389	8-14
-	556	28-36
H_2O	413	8-14
-	662	28-36

Water + Ferric nitrate (FeN_3O_9)

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.0748	-0.478	0.7480	-6.735
.1050	-0.667	1.0472	-11.433
.1496	-0.952	1.3464	-17.260
.2992	-2.076	1.4960	-21.400
.4488	-3.426		

M	d	M	d
	0°		
0.0748	1.016704	0.7480	1.122188
.1496	.031484	1.0472	.193728
.2992	.059356	.3464	.247724
.4488	.082232	.4960	.273804

Franz, 1872

%	d	%	d
	17.5°		
0	0.9997	33	1.2929
1	1.0067	34	.3037
2	.0147	35	.3147
3	.0227	36	.3263
4	.0307	37	.3379
5	.0385	38	.3474
6	.0459	39	.3610
7	.0532	40	.3728
8	.0606	41	.3846
9	.0680	42	.3964
10	.0756	43	.4082
11	.0838	44	.4200
12	.0920	45	.4320
13	.1002	46	.4447
14	.1084	47	.4573
15	.1168	48	.4700
16	.1254	49	.4827
17	.1340	50	.4953
18	.1426	51	.5103
19	.1510	52	.5253
20	.1596	53	.5403
21	.1696	54	.5552
22	.1796	55	.5702
23	.1896	56	.5872
24	.1996	57	.6042
25	.2094	58	.6212
26	.2196	59	.6382
27	.2298	60	.6551
28	.2400	61	.6743
29	.2501	62	.6935
30	.2605	63	.7127
31	.2713	64	.7319
32	.2821	65	.7510

Miescher, 1930

c	d
	20°
10.7	1.0792
26.5	.1964
52.0	.3770

Jones, 1904; Jones and Getman, 1904

M	n_D	M	n_D
	0°		
0.0748	1.32869	0.7480	1.35797
.1496	.33191	1.0472	.37075
.2992	.33395	.3464	.37693
.4488	.34536	.4960	.38873

M	molar. conduct.	M	molar. conduct.
	0°		
0.0748	152.5	0.7480	82.0
.1496	136.4	1.0472	63.4
.2992	117.2	.3464	48.2
.4488	102.0	.4960	43.2

Miescher, 1930

w.l. Å	Verdet's constant . 10^5 10.7 g/100cc	26.5 g/100cc	52.0 g/100cc
	20-21°		
6830	914	686	306
6320	973	726	316
6120	1037	765	316
5920	1103	812	320
5815	1145	833	316
5715	1188	862	316
5615	1229	887	310
5515	1270	909	304
5410	1325	943	300
5310	1377	969	296
5210	1428	999	284
5105	1478	1028	280
5005	1543	1065	-
4905	1600	1111	-
4805	1675	1166	-
4705	1740	-	-

Water + Ferric perchlorate ($FeCl_3O_{12}$)

Lindstrand, 1936

%	f.t.	%	f.t.
(10+1)		(9+1)	
74.32	0	83.76	40
78.64	20	85.50	50
79.86	25	87.05	55
80.84	30	88.53	60
81.74	35		
82.70	40		
83.19	42		

%	t	d	%	t	d
	(10+1)			(9+1)	
74.32	0	1.613	83.76	40	1.693
78.64	20	.649	85.50	50	.700
79.86	25	.656	87.05	55	.707
80.84	30	.666	88.53	60	.714
81.74	35	.676			
82.70	40	.684			
83.19	42	.688			

Water + Ferric sulfate ($Fe_2O_{12}S_3$)

Franz, 1872

%	d	%	d
	17.5°		
0	0.9987	35	1.3764
5	1.0413	40	.4487
10	.0840	45	.5279
15	.1310	50	.6128
20	.1810	55	.7029
25	.2410	60	.7984
30	.3073		

Hager, 1876

%	d	%	d	%	d
		18°			
0	0.999	15	1.149	30	1.335
1	1.007	16	.160	31	.349
2	.016	17	.171	32	.363
3	.026	18	.182	33	.378
4	.035	19	.194	34	.393
5	.045	20	.206	35	.409
6	.056	21	.218	36	.425
7	.066	22	.230	37	.440
8	.076	23	.243	38	.456
9	.086	24	.256	39	.472
10	.096	25	.269	40	.488
11	.105	26	.282	41	.504
12	.116	27	.295	42	.521
13	.127	28	.308	43	.538
14	.138	29	.321	44	.555

Gerlach, 1889

%	d	%	d
	15°		
0	0.999	30	1.330
10	1.095	40	1.477
20	1.204	50	1.649

Mc Gregor, 1894

%	d
15°	
10	1.0951
20	1.2040

Cabrera and Moles, 1913

%	d	%	d
	20°		
0	0.9971	3.978	1.0340
0.351	1.0013	5.632	.0501
0.469	.0015	6.678	.0612
1.173	.0080	7.986	.0730
2.326	.0185	11.259	.1132
2.963	.0237	14.099	.1140
3.401	.0300	15.382	.1525

%	χ	%	χ
	20°		
0	-0.720	3.978	+1.7465
0.3508	-0.5132	6.678	3.4541
0.4668	-0.4439	7.986	4.2446
1.173	-0.0137	11.260	6.4507
2.326	+0.6989	14.050	8.1863
2.963	+1.1066	15.380	9.0020
3.380	+1.3822		

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
26.390	1.2240
52.781	1.4319

Miescher, 1930

c	d
18°	
15.4	1.1354
21.8	.1905
48.0	.4059

w.l.	Verdet's constant . 10^5		
	15.4 g/cc	21.8 g/cc	48.0 g/cc
6530	809	-	254
6420	-	723	250
6320	860	-	246
6220	-	761	240
6120	908	-	240
6020	-	805	236
5920	967	825	226
5815	-	850	210
5715	1034	875	200
5615	-	903	184
5515	1096	922	165
5410	-	955	145
5310	1177	983	121
5210	-	1011	89
5105	1256	1037	48
5005	-	1068	4
4905	1355	1094	-52
4805	-	1120	-
4705	1439	-	-
4505	1542	-	-

Water + Ferric ammonium sulfate ($FeNH_4O_4S_2$)

Gerlach, 1889

%	d	%	d
15°			
0	0.999	16.009	1.121
2.838	1.022	19.846	.147
5.843	.046	23.946	.174
9.027	.070	28.337	.202
12.410	.095		

Water + Chromic chloride ($CrCl_3$)

Smith, 1947

Isopiestic solutions.

m_1	m_2	m_1	m_2
0.1046	0.186	0.829	2.173
0.3203	0.633	1.087	3.174
0.590	1.359	1.322	4.111
0.672	1.629		

m_1 = m of $CrCl_3$. m_2 = m of KCl .

Jones , 1904 ; Jones and Bassett, 1905

M	f.t.	M	f.t.
0.05	-0.268	0.7	-4.55
.10	-0.510	0.9	-6.30
.20	-1.030	1.0	-7.46
.30	-1.570	1.5	-15.00
.40	-2.160	2.0	-27.00
.50	-2.910	2.25	-33.00
.60	-3.610		

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
25°			
0.1	0.811	0.7	1.111
.2	.833	.8	.181
.3	.875	.9	.250
.4	.926	1.0	.319
.5	.983	.2	.443
.6	1.045		

Jones, 1904 ; Jones and Bassett, 1905

M	d	M	d
0°			
0	0.999868	0.7	1.088928
0.05	1.009916	0.9	.111756
.1	.014208	1.0	.122736
.2	.025832	1.5	.180612
.3	.036884	2.0	.239728
.4	.050704	2.25	.263136
.5	.062772		
.6	.077340		

Next Page

Heydweiller, 1921

N	d
18°	
0.2	1.00950
0.5	.02349
1	.04607
2	.09030
3	.13320

Liquier-Milward, 1935

t	d	t	d
7.90 %			
7.4	1.084	35.3	1.073
9.8	.083	57.6	.062
13.8	.082	68.8	.057
14.0	.082	70.2	.056
11.03 %			
8.9	1.123	41.8	1.110
9.8	.123	50.3	.106
10.4	.122	55.3	.103
18.4	.119	58.3	.101
25.9	.116	61.2	.099
33.8	.114	63.7	.097
18.42 %			
8.4	1.239	43.7	1.223
14.9	.2365	46.5	.221
15.3	.2365	51.0	.218
16.1	.236	58.3	.213
16.2	.236	58.8	.213
17.6	.235	63.4	.210
19.8	.234	68.1	.207
19.9	.234	69.9	.205
21.6	.233	73.2	.203
25.6	.232	74.8	.202
28.8	.231	77.2	.200
27.28 %			
10.5	1.453	43.6	1.437
10.8	.453	49.4	.433
19.2	.449	51.9	.432
34.2	.441	66.1	.424

Partington and Tweedy, 1926

%	η (water=1)	%	η (water=1)
25°		18°	
green			
3.58	1.0635	4.05	1.0548
6.35	.1111	6.86	.0890
12.91	.1984	10.04	.1166
16.34	.2462	12.79	.1507
21.50	.3333	17.07	.2329
35.67	.5397	22.38	.2808
50.32	.8095	24.80	.3493
		36.36	.5616
		52.40	.8904
violet			
3.68	1.0793	1.63	1.0342
5.75	.1270	3.46	.0753
8.48	.1944	5.48	.1164
11.23	.2381	8.02	.1849
16.27	.3730	9.44	.2123
17.58	.3968	13.57	.3356
20.61	.4841	16.81	.4315
33.64	.9206	26.82	.6575
62.20	2.3571	40.18	2.1644

Jones, 1904 and Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
0°			
0.05	100.20	0.7	55.47
0.10	89.14	0.9	48.62
0.20	79.70	1.0	45.40
0.30	72.37	1.5	32.19
0.40	68.11	2.0	22.60
0.50	63.12	2.25	14.46
0.60	59.27	2.273	18.31

Heydweiller, 1921

N	λ
18°	
0.2	75.5
0.5	68.6
1	56.8
2	44.8
3	35.2

Previous Page

Liquier-Milward, 1935

t	χ	t	χ
	7.90 %		
7.4	9.06	57.6	7.77
13.8	8.89	68.8	7.55
14.0	8.89	70.2	7.58
35.3	8.34		
	11.03 %		
8.9	9.08	50.3	8.00
10.4	9.03	55.3	7.83
18.4	8.71	58.3	7.80
25.9	8.51	61.2	7.74
33.8	8.31	63.7	7.70
41.8	8.23		
	18.42 %		
8.4	8.86	46.5	7.96
14.9	8.78	51.0	7.85
15.3	8.78	58.3	7.71
16.1	8.71	58.8	7.69
16.2	8.70	63.4	7.52
17.6	8.67	68.1	7.47
19.9	8.64	69.9	7.48
21.6	8.58	74.8	7.41
25.6	8.43	77.2	7.37
28.8	8.40	73.2	7.32 sic.
43.7	7.97		
	27.28 %		
10.5	8.62	49.4	7.71
19.2	8.42	51.9	7.70
34.2	8.01	66.1	7.38
43.6	7.88		

Elias, 1918

%	χ_{mol}	d^{25}
	green	
1.724	6181	-
3.343	"	-
9.503	6179	-
19.65	"	1.181-1.192
31.14	"	1.319-1.325
	violet	
1.725	6000	-
4.366	"	-
19.14	"	1.198
21.52	"	1.230

Water + Chromic bromide ($CrBr_3$)

Heydweiller, 1921

N	d	λ
	18°	
0.1	1.00839	87.7
0.2	.0170	81.7
0.5	.04213	72.8
1	.0837	63.8
2	.1661	51.3
3	.2474	41.3
4	.3278	33.22

Water + Chromic nitrate (CrN_3O_9)

Smith, 1947

m_1	m_2	m_1	m_2
0.0564	0.0966	0.734	1.740
.1202	.2098	0.926	2.380
.2208	.4043	1.164	3.263
.3485	.6850	1.443	4.385
.5520	1.2020		

m_1 = molality of CrN_3O_9 m_2= molality of KCl

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.795	0.7	1.055
.2	.818	.8	.111
.3	.860	.9	.168
.4	.906	1.0	.227
.5	.953	.2	.343
.6	1.003	.4	.456

Jones, 1904; Jones and Getman, 1904

M	f.t.	M	f.t.
0.0467	-0.280	0.9340	-8.800
.0934	-0.553	1.1208	-11.570
.1868	-1.143	.3076	-14.670
.3736	-2.493	.4944	-19.140
.5604	-4.153	.8680	-29.500

Traube, 1895

%	d
15°	
0	0.99913
3.389	1.02699
7.550	.06252
16.536	.14602
29.082	.28163

Jones, 1904; Jones and Getman, 1904

M	d	M	d
	0°		
0.0934	1.017200	1.1208	1.202672
.1868	.035312	.3076	.233856
.3736	.069208	.4944	.265308
.5604	.102096	.8680	.333516
.9340	.172288		

Heydweiller, 1921

N	d	N	d
	18°		
0.1	1.00636	2	1.1228
.2	.01256	3	.1823
.5	.03119	4	.2411
1.0	.0620	5	.2994

Partington and Tweedy, 1926

%	η	%	η
	18°		25°
3.55	1102	1.91	934
4.99	1131	3.18	953
7.16	1172	3.58	963
9.56	1221	5.93	999
15.07	1373	6.05	1006
19.73	1523	8.99	1050
23.62	1669	10.98	1093
26.58	1823	12.61	1122
29.11	1948	18.58	1281
		25.37	1498

Jones, 1904; Jones and Getman, 1904

M	molec.conduc.	M	molec.conduc.
	0°		
0.0934	143.1	1.1208	61.14
.1868	130.6	.3076	50.85
.3736	108.9	.4944	42.74
.5604	96.4	.8680	29.65
.9340	73.0		

Heydweiller, 1921

N	λ	N	λ
	18°		
0.1	80.3	2	43.25
.2	73.6	3	33.34
.5	64.9	4	25.54
1.0	55.9	5	18.97

Water + Chromic sulfate ($Cr_2O_{12}S_3$)

Smith, 1947

m_1	m_2	m_1	m_2
0.1144	0.1270	0.634	0.962
.1965	.2160	.779	1.385
.2143	.2379	.878	1.764
.3775	.4524	1.029	2.396
.5650	.7960	1.304	3.880

m_1= molality of $Cr_2O_{12}S_3$ m_2= molality of KCl

Robinson and Stokes, 1949

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.1	0.414	0.7	0.585
.2	.401	.8	.657
.3	.412	.9	.740
.4	.437	1.0	.832
.5	.473	1.2	1.031
.6	.524		

Gerlach, 1889

% violet	d	% green	d
	(18+1)	15°	
0	0.9991		
1.50	1.0266	0	0.9991
3.00	.0551	3.00	1.0501
6.00	.1140	6.00	.1060
9.00	.1775	9.00	.1670
12.00	.2469	12.00	.2329
15.00	.3238	15.00	.3044
		18.00	.3813
		21.00	.7637
		24.00	.5516

Manchot, Jahrstorfer and Zepter, 1924

c	d
25°	
22.356	1.1637
44.712	1.3280

Water + Chromic ammonium sulfate ($CrNH_4O_8S_2$)

Smith, 1947

m_1	m_2	m_1	m_2
0.0485	0.0608	0.3695	0.3996
.1106	.1297	.3829	.4125
.1196	.1381	.4148	.4458
.1223	.1426	.4485	.4806
.1424	.1643	.4608	.4947
.1556	.1778	.4855	.5230
.2004	.2243	.5048	.5464
.2077	.2324	.5472	.5911
.2225	.2465	.5558	.6000
.2422	.2677	.5639	.6090
.2866	.3141	.5692	.6180
.2983	.3250	.5808	.6320
.3329	.3607	.5957	.6470
.3483	.3774	.6807	.7450
.3535	.3817		

m_1 = molality of $CrNH_4O_8S_2$

m_2 = molality of KCl

Gerlach, 1889

% violet	d	% green	d
	15°		
0	0.9991	0	0.999
2.195	1.0191	5.806	1.043
4.390	.0396	12.327	.090
6.585	.0601	19.704	.141
sat.sol.	.0691	28.121	.196
		37.810	.254
		49.083	.316
		62.364	.383
		78.243	.455
		97.562	.531

Water + Potassium chromium sulfate ($KCrS_2O_8$)

Lescoeur and Mathurin, 1888

t	p	
	sat.sol.(green form)	
63	127	
85	245	
100	520	

t	p dissoc.	
10	3.7	
15	6.0	
20	9.1	
25	12.3	(12+1)
30	16.7	
35	23.8	

Smith, 1947

Isopiestic solutions

m_1	m_2	m_1	m_2
		25°	
0.1101	0.0974	0.492	0.471
.229	.207	.501	.481
.243	.220	.523	.502
.260	.238	.668	.644
.373	.351	.791	.751
.405	.386	1.233	1.095

m_1= molality of KCl

m_2= molality of $KCrS_2O_8$

Gerlach, 1889

%	d	%	d
	15° (12+1)		
violet form		green form	
0	0.99913	5.677	1.049
2.839	1.02636	11.355	1.102
5.677	1.05408	17.032	1.160
8.516	1.08295	22.710	1.224
sat.sol.	1.09754	28.387	1.294
		34.065	1.370
		39.742	1.452
		45.420	1.540
		51.097	1.634

Franz, 1872

%	d	%	d	%	d
		17.5°			
0	0.9987	24.	1.0940	48	1.2659
1	1.0022	25	.0990	49	.2767
2	.0057	26	.1046	50	.2777
3	.0092	27	.1102	51	.3039
4	.0127	28	.1158	52	.3201
5	.0161	29	.1214	53	.3363
6	.0196	30	.1260	54	.3524
7	.0231	31	.1320	55	.3686
8	.0266	32	.1380	56	.3858
9	.0301	33	.1440	57	.4030
10	.0329	34	.1499	58	.4202
11	.0365	35	.1557	59	.4374
12	.0401	36	.1622	60	.4547
13	.0437	37	.1686	61	.4724
14	.0473	38	.1751	62	.4901
15	.0510	39	.1816	63	.5078
16	.0554	40	.1880	64	.5255
17	.0598	41	.1971	65	.5433
18	.0642	42	.2062	66	.5614
19	.0684	43	.2153	67	.5796
20	.0732	44	.2244	68	.5978
21	.0784	45	.2336	69	.6160
22	.0836	46	.2444	70	.6342
23	.0888	47	.2551		

Ferrero, 1901

%	d
20.5°	
13.75	1.04750
20.5	1.07611

Rakuzin and Rosenfeld, 1927

%	d 1st sample	d 2nd sample
	20°	
4.8	1.0378	1.0199
9.1	.0641	.0402
16.7	.1076	.0793
23.1	.1232	-
28.6	.1623	.1564
33.3	.2134	-
37.5	.2531	-
41.2	.2764	-
44.4	.3524	-
47.4	.4062	-
50.0	.4404	-
52.4	.5237	-
53.31(satd)	.6683	-

D'Arcy, 1896

t	η (water^{15}=1)	t	η (water^{15}=1)
	0.7 N		
violet		green	
16	1.245	16.05	1.132
16.2	.224	29.85	0:826
29.8	0.896	49.35	0.579
49.3	.620	69.4	0.452
58.8	.528		

Ferrero, 1901

t	η	
20.5 %		
20.5	141.68	violet
52.5	141.19	
58.5	139.92	change of colour
60.2	137.41	green
63.0	135.31	
66.2	134.13	
70.0	132.37	
75.0	130.70	
80.5	129.46	
83.0	129.43	
85.0	129.45	
13.75 %		
20.5	126.09	violet
52.5	126.19	
55.0	123.84	change of colour
60.0	120.99	green
65.0	119.82	
70.0	118.93	
75.2	118.84	
81.2	118.54	
85.0	118.54	

Rehberg, 1949

Transparency of solutions

Water + Uranium sulfate ($U_2O_{12}S_3$)

Pareau, 1877

t	p dissoc. (4+1)-(1+1)	t	p dissoc. (1+1)- anh.
24.1	7.7	15.7	4.9
30.2	11.3	33.6	12.8
37.1	16.7	41.5	19.2
46.1	29.0	45.4	23.5
50.8	32.7	49.6	28.9

Water + Ceric ammonium nitrate ($CeH_8N_8O_{18}$)

Wolff, 1905

%	f.t.
58.49	25.0
61.79	35.2
64.51	45.3
66.64	64.5
69.40	85.6
88.03	122

Water + Cobalt chloride pentammine ($CoCl_3H_{15}N_5$)

Kurnakow, 1895

%	f.t.
14.09	0
19.73	16.2
19.92	16.9

Water + Lithium thioantimonate (Li_3SbS_4)

Donk, 1908

%	f.t.	%	f.t.
7.1	-1.7	35.3	-26.2
12.8	-3.2	40.4 E	-42.0
15.4	-4.2	45.5 (10+1)	0
17.5	-5.1	46.9	+10
20.0	-6.3	50.1	30
23.2	-10.8	51.3	50
28.5	-15.9		

Water + Sodium thioantimonite (Na_3SbS_3)

Sklyarenko and Kaplan, 1950

M	κ	M	κ
	20°		
0.6545	935	0.1022	214
0.4990	805	0.0828	174
0.4055	687	0.0631	143
0.3284	578	0.0513	117
0.2505	466	0.0425	101
0.2036	390	0.0188	494
0.1649	325	0.00945	258
0.1257	260		
	25°		
0.5621	978	0.0826	191
0.5089	904	0.0714	170
0.4221	785	0.0664	159
0.3281	654	0.0556	143
0.2830	579	0.0414	108
0.2561	524	0.0283	0769
0.2121	452	0.0216	0591
0.1643	364	0.0143	0415
0.1422	319	0.0072	0221
0.1061	240		
	60°		
1.321	269	0.2505	929
1.090	254	0.2036	777
0.9940	243	0.1649	649
0.8279	221	0.1257	517
0.6545	190	0.0828	346
0.4990	159	0.0513	249
0.4055	137	0.0425	209
0.3284	114	0.0188	0985

Water + Lithium fluogermanate (Li_2GeF_6)

Müller, 1926

53.92 g/100 cc f.t. = 25°

Water + Thallium fluogermanate (Tl_2GeF_6)

Müller, 1926

c	f.t.
14.13	0
34.58	30

Water + Silver fluogermanate (Ag_2GeF_6)

Müller, 1926

88.03 g/100 cc f.t. = 30°

Water + Sodium germanate (Na_2GeO_3)

Pugh, 1934

%	f.t.	%	f.t.
2.12	-0.7	21.74	30
5.00	-1.5	24.48	35
7.41	-2.0	29.58	45
9.15	-2.4	34.43	52.5
11.00	-3.0	41.86	63.5
12.80	0	46.20	70.5
17.67	+15.2	49.80	75.5
18.90	20.2	56.71	83.5
20.52	25		

%	t	d
17.67	15.2	1.1992
18.90	20.2	.2168
20.52	25	.2344
21.74	30	.2590

Water + Platinum chloride ($PtCl_4$)

Kurnakov and Nikitina, 1940

%	f.t.	%	f.t.
39.99	0	74.66	70
58.70	25	78.60	80
62.38	40	82.50	90
69.03	50	85.11	98
74.00	60		

Precht, 1879

%	d	%	d
	at room temp.		
1	1.009	26	1.300
2	.018	27	.315
3	.027	28	.330
4	.036	29	.340
5	.046	30	.362
6	.056	31	.387
7	.066	32	.395
8	.076	33	.413
9	.086	34	.431
10	.097	35	.450
11	.108	36	.469
12	.119	37	.488
13	.130	38	.500
14	.141	39	.523
15	.153	40	.546
16	.165	41	.568
17	.176	42	.591
18	.188	43	.615
19	.201	44	.641
20	.214	45	.666
21	.227	46	.688
22	.242	47	.712
23	.256	48	.736
24	.270	49	.760
25	.285	50	.785

Kurnakov and Nikitina, 1940

%	d	η(water=1)	%	d	η(water=1)
	yellow		25°	red	
8.89	1.0619	1.1046	8.32	1.0856	1.5413
12.01	.1180	.2310	13.10	.1199	.7500
15.03	.1508	.2902	20.25	.2227	.9100
18.05	.1823	.3435	24.00	.2811	2.1001
20.30	.1942	.3606	28.18	.3249	.4000
24.08	.2266	.4008	31.87	.3687	.6300
27.99	.2616	.4385	34.37	.4111	.7660
33.83	.3724 sic	.4721	38.07	.4695	3.1000
37.11	.2924	.5300	40.51	.4987	.3340
43.23	.6124	.6225	43.27	.5417	.6005
48.34	.6286	.6786	45.09	.5702	.8009
50.73	.7231	.7600	46.25	.5748	.9983
			47.99	.6724	4.2000
			49.98	.8224	.4512
			54.43	.8500	.8784

Water + Tris-ethylendiamine-Platinum chloride ($C_6H_{24}N_6Cl_4Pt$)

Brubaker, 1956

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.05582	0.5797	0.2355	0.4848
.08505	.5408	.2705	.4802
.12030	.5184	.3034	.4833
.18110	.4991	.3073	.4814
.22310	.4882	.3281	.4774

Water (H_2O) + Sodium chloroplatinate (Na_2PtCl_6)

Kurnakov, 1895

%	t	d	n Li	n Na
29.123	18.8	1.28259	1.38749	1.39085

Kurnakov and Nikitina, 1940

wt %	mol %	f.t.
38.70	2.44	0
44.91	3.13	25
47.87	3.51	30
49.04	3.67	35
50.67	3.79	40
57.07	5.01	60
59.12	5.42	70
65.25	6.93	80
68.10	7.81	90
71.91	9.21	98

Water + Stannic chloride ($SnCl_4$)

Gerlach, 1865

%	d	%	d	%	d
		15°			
0.74	1.005	24.50	1.217	48.26	1.5242
1.48	.011	25.25	.2257	49.00	.537
2.22	.017	25.98	.231	49.75	.549
2.96	.023	26.72	.241	50.49	.562
3.71	.029	27.46	.249	51.24	.574
4.45	.035	28.20	.258	51.97	.5859
5.19	.041	28.95	.266	52.71	.6000
5.93	.047	29.70	.2744	53.45	.613
6.67	.052	30.44	.283	54.19	.626
7.42	.058	31.18	.292	55.94	.640
8.16	.065	31.92	.301	55.68	.6529
8.90	.071	32.66	.309	56.42	.668
9.64	.077	33.41	.3181	57.16	.682
10.38	.083	34.15	.328	57.90	.697
11.13	.089	35.89	.337	58.65	.710
11.87	.096	36.63	.346	59.40	.7256
12.92	.103	37.37	.356	60.14	.741
13.96	.109	38.12	.3649	60.88	.757
14.15	.110	37.86	.375	61.62	.773
14.85	.125	38.61	.385	62.36	.789
15.59	.129	39.35	.395	63.11	.7951
16.34	.136	40.09	.405	63.85	.822
17.08	.143	40.83	.4142	64.60	.840
17.82	.150	41.57	.425	65.34	.857
18.56	.1571	42.32	.436	66.08	.874
19.30	.164	43.06	.446	66.81	.8920
20.05	.172	43.80	.457	67.55	.911
20.79	.179	44.55	.4671	68.30	.930
21.54	.186	45.29	.479	69.04	.945
22.27	.193	46.04	.490	69.78	.967
23.01	.201	46.78	.502	70.52	.9864
23.76	.209	47.53	.513		

Heermann, 1907

%	d	%	d
	17.5°		
10.91	1.2078	23.11	1.5611
11.35	.2181	23.56	.5782
11.79	.2284	24.02	.5956
12.23	.2390	24.47	.6136
12.60	.2497	24.93	.6320
13.11	.2606	25.38	.6505
13.56	.2718	25.84	.6696
14.00	.2831	26.30	.6892
14.45	.2947	26.77	.7092
14.90	.3074	27.24	.7298
20.38	.4658	27.70	.7507
20.83	.4808	28.17	.7723
21.29	.4963	28.64	.7944
21.74	.5120	29.12	.8170
22.20	.5280	29.45	.8263
22.65	.5445		

Chèneveau, 1907

%	d	%	d
22°			
0	0.9978	21.34	1.1415
7.74	1.0492	27.35	1.1861
14.83	1.0951	32.90	1.2312
18.15	1.1187		

%	n C	n D	n Tl	n F	n G'
18.6°					
20.67	1.35881	1.36094	1.36318	1.36583	1.36972

Heydweiller, 1921

N	d	N	d
18°			
0.5	1.02553	3	1.1449
1	1.05023	4	1.1910
2	1.0979		

Guillaume, 1946

%	d	%	d
20°			
16.8	1.139	60.5	1.7510
37.4	1.3664	100	2.216
54.0	1.6229		

Chèneveau, 1907

%	n_D	%	n_D
22°			
0	1.3328	21.34	1.3610
7.74	1.3434	27.35	1.3694
14.83	1.3522	32.90	1.3781
18.15	1.3566		

Heydweiller, 1921

N	λ	N	λ
18°			
0.5	216.8	3	47.9
1	121.7	4	32.7
2	66.9		

Guillaume, 1946

%	n 5780 Å	(α) *magn. 10^6
20°		
16.8	1.3631	4.221
37.4	.4047	.553
54.0	.4517	.866
60.5	.4774	5.008
100	.5162	5.872

* in radians, gauss, centim.

Water + Stannic bromide ($SnBr_4$)

Pickering, 1895

%	f.t.	%	f.t.
26.97	-7.5	71.24	17.17
42.47	-16.0	71.26	17.58
49.74	-20.0	71.97	18.41
56.36	-33.5	72.26	18.20
66.70	+9.62(8+1)	72.53	18.51
67.12	10.13	72.72	18.41
67.82	12.51	73.63	18.90
68.47	14.99	100	29.36
70.97	16.85		

Water (H_2O) + Sodium stannate (Na_2SnO_3)

Traube, 1895

%	d	%	d
20°			
0	0.9982	10.78	1.095
4.73	1.039	19.63	1.183

Water (H_2O) + Potassium stannate (K_2SnO_3)

Traube, 1895

%	d
20°	
0	0.998
22.73	1.212
36.53	1.379

Water + Diammonium fluozirconate ($ZrH_8F_6N_2$)

van Hevesy,Christiansen and Berglund, 1925

M	f.t.	$d_{20°}$
0.611	0	-
1.050	20	1.154
1.842	45	-
2.960	90	-

Water + Triammonium fluorzirconate ($ZrH_{12}F_7N_3$)

van Hevesy,Christiansen and Berglund, 1925

M	f.t.	$d_{20°}$
0.360	0	-
0.551	20	1.086
0.788	45	-

Water + Diammonium fluohafniate ($HfH_8N_2F_6$)

van Hevesy, Christiansen and Berlund, 1925

M	f.t.
0.890	0
1.425	20

Water + Triammonium fluohafniate ($HfH_{12}N_3F_7$)

van Hevesy, Christiansen and Berglund, 1925

M	f.t.
0.425	0
0.588	20

Water + Thorium chloride ($ThCl_4$)

Robinson, 1955

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.05	0.731	0.7	1.129
.1	.736	.8	.214
.2	.776	.9	.302
.3	.840	1.0	.390
.4	.906	.2	.536
.5	.974	.4	.665
.6	1.048	.6	.847

Heydweiller, 1921

N	d	λ
	18°	
0.5	1.0436	61.0
1	.0868	54.0
2	.1720	44.33
3	.2558	36.3
4	.3388	29.81

Jauch, 1921

N	U
18°	
0.5	0.9443
1	.8954
2	.8118
3	.7435
4	.6860

Water + Thorium nitrate ($ThO_{12}N_4$)

Fricke, 1929

mol%	p 0°	p 10°
4.09	3.408	9.533
4.48	3.265	9.165
4.93	3.085	8.738
5.50	2.882	8.210
6.07	2.698	7.724

Robinson and Stoke, 1949

m	osmotic coefficient	m	osmotic coefficient
		25°	
0.1	0.675	1.4	1.120
0.2	0.685	1.6	1.192
0.3	0.705	1.8	1.259
0.4	0.734	2.0	1.325
0.5	0.770	2.5	1.455
0.6	0.807	3.0	1.546
0.7	0.846	3.5	1.616
0.8	0.885	4.0	1.659
0:9	0.925	4.5	1.688
1.0	0.965	5.0	1.706
1.2	1.044		

Misciattelli, 1930

%	f.t.	%	f.t.	
1	-0.2	47.7	-15	
2	-0.5	49.8	-16.6	
5.2	-1	51	-19.1	
9	-1.5	53.1	-23.3	
13	-2.1	55.7	-25	
16	-2.9	58	-28.6	
20	-4	59.2	-31.3	
23.5	-4.6	60.6	-35	
26.4	-5.4	62	-40.6	
27.3	-5.6	64	-43.5	(6+1)
33	-6	64.2	-22	
37	-6.6	65	0	
41	-9	65.2	+10	
43	-11.2	65.6	20	
44.5	-12.2			
46.2	-13.5			

Marshall, Gill and Secoy, 1951

%		f.t.
67.07	(6+1)	37.3
69.78		54.5
73.39		72.0
76.39		90.2
78.56		99.7
81.11		110.4
81.50		110.9
-	(6+1)-(4+1)	111
82.01		120.6
82.41		128
82.85		130.5
84.27		139.5
85.30		146.0
85.81		149.0
-	(4+1)-(x+1)	151
87.41		159
91.82		211

Koppel and Holtkamp, 1910

%	d	n_D
	15°	
0	0.9991	1.33337
0.53	1.0052	.33478
1.34	.0131	.33582
2.60	.0318	.33804
3.85	.0474	.34025
5.07	.0629	.34283
6.23	.0786	.34489
7.36	.0952	.34724
9.53	.1287	.35235
11.60	.1599	.35732
13.54	.1916	.36188

Water + Thorium sulfate (ThO_8S_2)

Dawson and Williams, 1900

t	κ	t	κ
(9+1)		(4+1)	
25.05	4.381	45.05	8.977
35.02	5.405	52.2	8.511
40.00	6.379	60.35	8.157
44.90	7.675		
49.80	9.372		

tr.t. (9+1) - (4+1) 47-48°

Water + Uranyl fluoride (UO_2F_2)

Johnson and Kraus, 1953

%	d 25°	d 30°
2.495	1.0202	-
4.985	1.0443	1.0429
5.025	1.0448	-
10.200	1.0983	1.0967
20.09	1.2163	1.2146
30.10	1.3632	-
30.18	1.3632	1.3607
40.30	1.5509	1.5481
50.27	1.7893	1.7859
61.63	2.1627	2.1589

%	n_D 25°	n_D 30°
1.002	1.33319	1.33266
2.495	.33417	.33366
4.985	.33586	.33531
5.025	.33597	-
7.481	.33764	1.33705
9.607	.33925	.33864
10.200	.33963	.33907
15.01	.34333	.34275
20.09	.34756	.34692
24.89	.34788	.34721
30.10	.35188	.35114
30.18	.35705	.35636
33.39	.36059	.35988
40.27	.36888	.36825
46.36	.37726	.37647
50.27	.38316	.38239
50.90	.38439	.38370
56.90	.39503	.39418
57.10	.39542	.39457
61.12	.40392	.40303
61.63	.40475	.40407

Brown, Bunger, Marshall and Secoy, 1954

N(25°)	λ 0°	λ 25°	λ 50°	λ 90°
0.0001	34.44	73.54	-	-
.0005	18.37	35.70	62.48	104.2
.0010	12.46	26.10	44.22	71.3
.0050	6.16	12.31	19.51	30.4
.01	4.16	9.17	14.41	22.7
.05	2.73	5.43	8.60	13.52
.1	2.37	4.74	7.50	11.88
.5	1.87	3.75	5.96	9.50
1.0	1.60	3.22	5.10	8.15
3.0	-	-	3.12	4.91
6.0	-	-	1.40	2.27

Water + Uranyl chloride (UO_2Cl_2)

Robinson and Lim, 1951

m_1	m_2	m_1	m_2
25°			
0.1047	0.1065	1.465	1.526
.1328	.1363	.706	.759
.2690	.2796	.777	.825
.3598	.3766	.797	.849
.3864	.4048	2.030	2.054
.4969	.5251	.084	.105
.6512	.6903	.427	.404
.8087	.8595	.442	.414
1.037	1.099	.765	.686
.288	.349	.818	.722
.344	.413	3.174	3.011

m_1 = m of UO_2Cl_2 m_2 = m of $CaCl_2$

m	d
25°	
0.5	1.1393
1.0	.2732
2.0	.5184
3.0	.7378
4.0	.9337

Kapustinskii and Lipilina, 1955

mol%	U	mol%	U
25°			
92.34	0.7557	36.07	0.8812
81.25	.7768	31.06	.8959
69.04	.8020	23.87	.9171
56.32	.8314	20.51	.9271
46.67	.8545	13.53	.9454

Water + Uranyl nitrate (UN_2O_8)

Tammann, 1885

%	p
100°	
12.65	746.9
25.72	725.1
31.69	710.7

Coulter, Pitzer and Latimer, 1940

m	p
25°	
3.11	17.2
1.98	18.7
1.257	21.8
0.578	23.0

Robinson and Lim, 1951

m_1	m_2	m_1	m_2	m_1	m_2
a		b		c	
0.9736	1.073	2.245	2.378	0.1010	0.1032
1.092	.202	.422	.538	.1691	.1755
.216	.335	.661	.734	.3034	.3213
.371	.498	3.127	3.071	.4879	.5245
.595	.733	.191	.117	.5484	.5920
.724	.869	.242	.152	.7737	.8431
2.008	2.151	.523	.329	1.041	1.145
.043	.188	.894	.534	.288	.408
.282	.408	4.306	.737	.522	.659
.356	.472	.481	.809	.823	.969
.473	.575	.810	.951	2.297	2.420
.574	.658	.971	4.021		
.641	.707	5.106	.067		
.664	.725	.319	.148		
.793	.837	.511	.235		
.942	.946				
.946	.936				
3.035	3.014				
.153	.097				
.401	.241				
.587	.263				

m_1 and m_2 - reasp. molalities of Uranyl nitrate and $CaCl_2$

preparations	a	b	c
pH at 0.2m	2.22	2.34	2.43

Vasilyev, 1910

%	f.t.	%	f.t.
10.83	-1.6	45.53	-12.1
12.24	-2.1	48.77	-2.2
17.19	-2.9	49.46	0
23.53	-4.4	51.67	+12.3
26.20	-6.0	57.17	+25.6
32.53	-7.9	61.26	+36.7
37.25	-11.2	65.12	+45.2
43.09	-18.1 E	67.76	+71.8

(6+1) f.t. = 60.2

(3+1) 121.5

(2+1) 179.3

Guempel, 1929

%	f.t.	%	f.t.
7.87	-0.3	55.90	25.0
17.33	-2.9	60.28	36.1
24.85	-5.4	64.20	43.6
26.20	-6.0	70.25	54.5
29.60	-7.3	71.95	56.1
43.04	-18.0	72.76	57.4
E	-18.1	74.13	58.2
46.25	-11.5 (6+1)	75.65	58.6 tr.t.
48.47	-5.55	76.83	62.0 (x+1)
45.27	-4.2	78.50	72.4
51.57	+7.6	80.20	80.9
52.80	15.0	81.13	88.5
54.40	20.0		

Benrath, 1942

%	f.t.	%	f.t.
79	80	86.7	138
80.6	94	87.9	156
82.9	108	90.5	175
84.6	115	91.6	187
85.8	116		

Marshall, Gill and Secoy, 1951

%	f.t.	%	f.t.
77.25	70 (3+1)	87.07	137
78.49	77.2	87.02	141.2
79.92	85	87.75	147
80.98	90.5	88.23	154.5
81.37	92	88.74	159
82.57	100	88.94	160
84.14	110	89.22	165.5
84.67	113 tr.t.	89.92	172
85.25	120 (2+1)	90.78	180
86.13	130	91.01	181
86.54	133	91.63	184

Oechsner and de Coninck, 1916

%	t	d	%	t	d
1	11.5	1.0045	9	16.5	1.0399
2	12.4	.0091	10	15.2	.0453
3	15.1	.0131	11	13.7	.0498
4	14.1	.0180	12	11.5	.0546
5	16.7	.0228	13	14.5	.0585
6	14.1	.0274	14	11.3	.0639
7	15.7	.0317	15	12.5	.0675
8	15.2	.0369	16	13.2	.0712

Germann, 1922

%	d	%	d
	24°		
5	1.035	30	1.302
10	.076	35	.376
15	.126	40	.457
20	.178	45	.540
25	.236	50	.626

Grant, Darch and al., 1948

M	d	M	d
	0°		
0.1529	1.0509	0.4842	1.1575
0.2291	.0740	0.6250	.1985
0.4127	.1328	0.8316	.2701
	20°		
0.2286	1.0714	0.6219	1.1920
0.4112	.1291	0.8271	.2634
0.4828	.1540		
	25°		
0.1523	1.0473	0.4821	1.1522
0.2283	.0703	0.6210	.1903
0.4107	.1275	0.8260	.2615
	30°		
0.1521	1.0459	0.4815	1.1506
0.2280	.0687	0.6201	.1886
0.4102	.1260	0.8247	.2596
	35°		
0.2277	1.0671	0.6191	1.1867
0.4094	.1240	0.8231	.2570
0.4804	.1488		
	40°		
0.1515	1.0420	0.4800	1.1477
0.2273	.0651	0.6181	.1847
0.4088	.1219	0.8215	.2546
	50°		
0.1507	1.0365	0.4777	1.1423
0.2262	.0609	0.6158	.1800
0.4071	.1175	0.8181	.2495
	60°		
0.1501	1.0324	0.4755	1.1369
0.2253	.0559	0.6131	.1748
0.4053	.1124	0.7145	.2440
	80°		
0.2230	1.0457	0.6074	1.1642
0.4019	.1026	0.8063	.2314
	100°		
0.2201	1.0315	0.5995	1.1492
0.3966	.0889	0.7980	.2187

Kapustinskii and Lipilina, 1956

%	d	%	d
	25°		
0.735	1.00316	24.92	1.25370
1.520	.00978	27.75	.29067
2.316	.01670	30.57	.33016
3.466	.02656	33.53	.37430
5.880	.04811	36.58	.42221
9.020	.07739	39.19	.46609
12.050	.10736	43.42	.54389
15.20	.14002	46.37	.60237
18.63	.17739	49.38	.66453
21.99	.21751	52.36	.73082

Grant, Darch and al., 1948

M	η		M	η
		0°		
0.1529	1900		0.4842	2260
0.2291	1980		0.6250	2430
0.4127	2190			
		20°		
0.2286	1110		0.4828	1280
0.4112	1220		0.6219	1330
		25°		
0.1523	943		0.4821	1140
0.2283	989		0.6210	1220
0.4107	1090			
		30°		
0.1521	844		0.4815	1020
0.2280	884		0.6201	1090
0.4102	974			
		35°		
0.2277	796		0.4804	915
0.4094	880		0.6191	978
		40°		
0.1515	691		0.4800	832
0.2273	726		0.6181	892
0.4088	794			
		50°		
0.1507	580		0.4777	696
0.2262	603		0.6158	745
0.4071	668			
		60°		
0.1501	492		0.4755	587
0.2253	514		0.6131	643
0.4053	568			

M	σ	M	σ
	0°		
0.2291	76.4	0.6250	76.8
0.4127	76.6	0.8316	77.6
	20°		
0.2286	73.3	0.6219	73.7
0.4112	73.4	0.8271	74.6
	25°		
0.2283	72.4	0.6210	73.0
0.4107	72.7	0.8260	73.9
	30°		
0.2280	71.6	0.6201	72.2
0.4102	72.0	0.8247	73.4
	35°		
0.2277	70.6	0.6191	71.5
0.4094	71.3	0.8231	72.7
	40°		
0.1515	69.8	0.4800	70.6
0.2273	70.4	0.6181	71.8
	50°		
0.1507	68.5	0.4777	69.2
0.2262	69.0	0.6158	70.3
	60°		
0.1501	66.9	0.4755	67.8
0.2253	67.5	0.6131	69.0
	80°		
0.2230	63.2	0.6074	64.2
0.4019	63.6	0.8063	65.0
	100°		
0.2201	59.9	0.5995	61.0
0.3966	60.6	0.7980	61.9

Kapustinskii and Lipilina, 1955

mol%	U	mol%	U
	25°		
4.02	0.9821	38.72	0.8669
15.98	.9394	50.91	.8387
22.47	.9178	69.54	.7914
25.47	.9081	89.04	.7495
28.62	.8960		

Nichols and Merritt, 1914

Absorption of the fluorescent bands at room temperature and at -185°

Water + Uranyl perchlorate ($UO_{10}Cl_2$)

Robinson and Lim, 1951

m_1	m_2	m_1	m_2
	25°		
0.1061	0.1141	2.591	3.667
0.1549	0.1714	2.749	3.907
0.2078	0.2333	2.869	4.090
0.2870	0.3324	3.289	4.721
0.3634	0.4293	3.687	5.327
0.3806	0.4520	3.907	5.685
0.5733	0.7109	4.242	6.267
0.7627	0.9773	4.601	6.949
0.9657	1.268	4.853	7.480
1.329	1.798	5.039	7.897
1.602	2.200	5.179	8.225
1.859	2.580	5.458	8.930
1.992	2.781		

m_1 = m of $UO_{10}Cl_2$ m_2 = m of $CaCl_2$

Water + Uranyl sulfate (UO_6S)

Tammann, 1885

%	p	%	p
	100°		
18.78	754.1	53.60	716.3
30.64	748.4	62.46	683.0
46.91	731.4	65.56	664.6

Robinson, 1952

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2	m_1	m_2
0.1132	0.06292	1.086	0.5895	4.605	4.226
0.1560	0.08389	1.251	0.7066	4.652	4.269
0.2265	0.1184	1.260	0.7129	5.317	5.022
0.2473	0.1257	1.346	0.7756	5.392	5.110
0.3261	0.1623	1.593	0.9711	5.983	5.726
0.3429	0.1703	1.789	1.137	6.069	5.806
0.3724	0.1849	2.036	1.363	6.371	6.122
0.4570	0.2284	2.386	1.705		
0.5796	0.2917	2.499	1.824		
0.6542	0.3322	2.776	2.117		
0.6648	0.3364	3.370	2.786		
0.7405	0.3806	3.373	2.893		
0.8799	0.4630	3.466	3.506		
1.0630	0.5776				

m_1 = m of UO_6S m_2 = m of ClNa

Secoy, 1948

m	f.t.	m	f.t.	
0.03	-0.0705	5.73	118.4	
.05	-0.1100	.66	119.9	
.10	-0.200	.90	125.2	
.30	-0.516	6.25	134.9	
.60	-0.948	.51	139.0	
2.87	-17.70	.55	140.7	
3.33	-22.90	.72	145.1	
3.82	-38.50 E	.89	149.6	
(3+1)		7.14	153.6	
		7.75	168.6	
3.84	-27.0	8.66	181	tr.t.
4.135	+30.0			
.171	35.0	(1+1)		
.225	40.0			
.31	46.8	8.97	152	
.50	57.3	7.75	287	
.73	71.7			
.70	80.8	(2+1)		
5.02	90.1			
.06	93.5	4.70	49.2	
.26	106.2	4.73	53.2	
.59	116.8	5.06	75.5	

Secoy, 1950

%	f.t.
(1+1)	
72.60	365
72.85	347
73.49	311

%	crit.t.	sat.t.
0.77	-	335.1
3.15	-	319.4
5.86	-	310.1
12.30	374.3	299.9
20.40	-	295.8
31.65	374.0	295.5
40.72	374.2	297.5
48.06	374.2	302.3
55.45	-	308.2
58.64	374.4	312.5
60.60	-	316.2
62.32	-	320.5
67.43	-	339.6
71.00	-	363.0

Orban, Barnett and al., 1956

M	d 20.0°	d 30.0°	d 44.8°
0.176	1.0539	1.0510	1.0445
.452	.1384	.1349	.1282
.671	.2050	.2012	.1939
.864	.2630	.2588	.2513
1.049	.3141	.3098	.3023
.321	.3911	.3865	.3779
.753	.5089	.5032	.4956
.983	.5672	.5613	.5516
2.865	.7797	.7736	-
2.858	-	-	1.7633
3.872	2.0059	1.9983	1.9856

M	d 59.8°	d 75.0°	d 90.0°
0.177	1.0343	-	-
.178	-	1.0293	-
.180	-	-	1.0193
.450	1.1208	-	-
.454	-	1.1113	-
.458	-	-	1.1008
.671	1.1862	-	-
.669	-	1.1761	-
.681	-	-	1.1654
.870	1.2432	-	-
.868	-	1.2334	-
.877	-	-	1.2219
1.049	1.2948	-	-
.053	-	1.2845	-
.055	-	-	1.2720
.323	1.3692	-	1.3458
.329	-	1.3579	-
.756	1.4867	-	-
.761	-	1.4735	-
.770	-	-	1.4620
.986	1.5430	-	-
2.002	-	1.5328	-
.009	-	-	1.5236
.870	1.7533	-	-
.880	1.9752	-	-
.959	-	1.9708	-
.999	-	-	1.9646

M	η 20.0°	η 30.0°	η 45.0°	η 60.0°	η 75.0°	η 90.0°
0.176	1120	881	659	516	422	349
.452	1370	1060	779	610	494	403
.671	1580	1230	893	685	547	460
.864	1800	1420	1010	779	608	501
1.049	2040	1580	1140	872	673	561
.321	2500	1930	1370	1040	836	656
.531	2910	2220	1570	1180	931	735
.753	3510	2640	1890	1400	1110	865
.978	4180	3150	2210	1670	1260	1010
2.865	8530	6210	4220	3070	2300	1820

Orban, Barnett and al., 1956

M	σ 20.0°	30.0°	45.0°	60.0°	75.0°
0.175	73.47	71.95	69.59	67.04	64.55
.446	74.00	72.56	69.98	67.43	64.83
.606	74.43	72.75	69.91	67.29	64.60
.728	74.18	72.58	70.23	67.75	65.06
.856	74.74	73.10	70.20	67.67	65.13
.881	74.12	73.22	70.68	68.50	65.87
.965	75.14	73.38	70.90	68.30	65.64
1.016	74.60	73.12	70.76	68.32	65.56
.184	75.83	74.13	71.44	68.96	66.26
.445	76.08	74.29	71.76	69.18	66.55
.672	76.09	74.26	72.18	69.61	66.68
.837	76.70	75.03	72.81	70.33	67.70
2.340	77.43	75.90	73.41	70.96	68.24

M	pH 24.7°	35.0°	44.8°	59.8°
0.00066	4.09	4.02	3.97	3.89
.0013	3.90	3.81	3.76	3.69
.0066	3.47	3.39	3.33	3.27
.0132	3.27	3.18	3.12	3.04
.0758	2.73	2.65	2.58	2.47
.157	2.48	2.38	2.29	2.18
.321	2.18	2.06	1.98	1.84
.824	1.72	1.61	1.52	1.38
1.731	1.25	1.13	1.01	0.89
3.117	0.68	0.52	0.41	0.30
3.850	0.33	0.18	0.04	0.12

Brown, Bunger and al., 1954

N	λ	N	λ	N	λ
0°		25°		50°	
0.0001008	83.13	0.0001	63.18	0.0000991	290.62
.000504	62.33	.0005	49.21	.0004955	198.79
.001008	54.05	.001	27.65	.000991	159.34
.002526	43.31	.1	22.18	.004955	85.92
.01263	26.29	.5	14.43	.00991	65.33
.02526	20.71	1.0	11.59	.04955	35.99
.0982	13.13	4.66	2.78	.0991	29.28
.3992*	9.37			.4955	20.53
1.0*	6.43			.991	17.03
3.075*	2.51			2.504*	9.59
6.065*	0.465			4.66*	4.69
				7.28*	1.02
90°		*125°		*200°	
0.0000968	521.60	0.0001	590.0	0.001	235.9
.000484	289.20	.0005	320.0	.005	124.8
.000968	227.23	.001	236.0	.01	105.4
.00484	107.42	.005	110.4	.05	66.0
.00968	79.53	.01	81.4	.10	55.0
.0484	45.24	.05	48.6	.50	39.8
.0968	38.53	.10	42.08	1.00	33.4
.484	28.64	.50	34.28	2.52	20.6
.968*	24.43	1.00	26.78		
2.520*	14.49	2.52	18.00		
4.660*	7.77	4.66	10.07		

* - concentrations on 25° basis.

Nichols and Merrits, 1914

Absorption of the fluorescent bands at -185°

Water + Uranyl acid sulfate ($UH_2O_{10}S_2$)

Tammann, 1885

%	p	t	p
	100°		
12.23	750.0	38.28	697.4
22.06	736.2	43.47	678.5
31.09	719.0	49.62	640.3

Water + Uranyl chromate ($UCrO_6$)

Loprest, Marshall and Secoy, 1955

%	t	%	t
Ice + L+V			
0	0	1.044	-0.06
0.099	-0.02	1.967	-0.09
.242	-0.03	2.452	-0.12
.496	-0.03		
(11+2) + L+V			
3.11	0	42.69	35.00
4.61	4.98	47.36	37.50
6.66	9.88	51.62	40.03
6.78	10.06	53.70	41.96
10.24	15.01	56.75	45.06
15.85	20.33	58.80	47.64
16.27	20.65	60.05	50.10
23.41	24.70	64.76	55.14
23.24	25.05	66.84	60.18
32.87	30.00		
(X+1) + L + V			
70.50	70.30	70.24	100.88
69.97	80.45	70.95	140
70.70	90.61		
(11+2) + (X+1) + L + V			
70.04	66.30		

Water + Chromic anhydride (CrO_3)

(see also chromic acid and bichromic acid)

Koppel and Blumenthal, 1907

%	b.t.	%	f.t.	%	f.t.
10.81	102	23.1	-6	61.54	0
24.08	104	28.6	-9.3	62.52	30
36.47	107	44.4	-24	65.12	60
45.15	110.5	50	-36		
54.56	116	54.5	-51		
61.54	120				
71.24	127				

Kremann, 1911

%	f.t.	%	f.t.
0	0.0	46.5	-36.3
8	-2.0	46.6	-37.4
13.9	-4.3	47.5	-38.9
20.0	-8.0	49.1	-43.0
28.5	-14.0	51.0	-48.4
30.2	-16.3	51.0	-48.0
31.1	-16.4	52.0	-52.0
34.1	-19.0	52.5	-54.3
35.4	-19.85	53.0	-60.0
35.0	-21.0	54.0	-64.0
38.0	-25.0	54.2	-66.0
38.7	-27.0	55.0	-74.0
39.8	-26.2	57.2	-105
40.8	-28.3	66.0	+82
43.4	-30.7	67.4	+100
44.0	-31.6	68.4	+115
46.0	-34.6		

Büchner and Prins, 1912

%	f.t.	%	f.t.
3.6	-0.90	53.30	-60
7.8	-1.90	61.70	-20
11.5	-3.70	62.24	0
14.1	-4.80	62.88	+24.8
24.9	-10.95	63.50	40
25.2	-11.70	64.83	65
33.5	-18.75	68.50	90
39.2	-25.25	70.70	122
49.1	-43.50	100	193-196

Rakovski and Tarasenkov, 1928

%	f.t.	%	f.t.
61.94	0	63.78	60
62.58	20	65.47	80
63.12	40	66.54	100

Vuillard, 1951 (fig.)

%	f.t.	E	%	f.t.	E
38	-22	-113	60	-102 tr.t.	-113
42	-24	"	62	+21	-102
52	-61	"	70	-	"
55	-75	"	80	-	"
57	-113	"	100	-	"

(7+2)

Gerlach, 1859

%	d	%	d
	17.5°		
0	0.999	35	1.310
5	1.036	40	.371
10	.075	45	.438
15	.117	50	.510
20	.160	55	.585
25	.206	60	.663
30	.256		

Zettnow, 1871

62.23% f.t. = 15°

%	t	d	%	t	d
62.23	15	1.7023	31.83	12	1.20714
37.8	19.2	.3448		20.1	.20264
	22	.34414		20.9	.20269
32.59	9.7	.22384	19.33	19	.15690
	15.2	.22100	12.39	19.5	.0957
	18.6	.21914	8.79	14.2	.0694
	35.0	.20940		18.0	.0679
			8.25	16.2	.0606
				17	.0600

%	d	%	d
	17.5°		
0	0.99871	19.33	1.15541
8.25	1.05839	37.80	.34446
8.79	.06667	62.23	.70010
12.34	.09429		

Wolf and Christofzik, 1955

M	η	M	η
	25°		
0.5	940	3.8	1400
0.9	950	4.0	1470
1.9	1600	5.0	1700
2.4	1140	5.3	1750
3.1	1280		

M	κ	M	κ
	25°		
0.3	1200	4.7	6500
1.0	2700	6.2	6200
1.6	4000	7.8	5000
2.2	5300	8.6	4500
3.5	6100		

Büchner and Prins, 1912

mol%	U	mol%	U
1.21	0.927	9.17	0.665
1.96	.876	16.61	.557
3.82	.803	21.98	.506

mol% initial	Q dil. (for 1.23mol%)	Q diss.
1.25	-	2467
2.00	16	2451
3.82	122	2345
9.00	469	1998
16.80	866	1601
19.90	1111	1356
20.90	1177	1290
22.10	1338	1129
23.20	1447	1020

Water + Chromic acid (CrO_4H_2)

(See also bichromic acid and chromic anhydrid)

Moore and Blum, 1930

M	d	M	d
	25°		
0.995	1.0731	6.001	1.3417
1.972	.1365	7.028	.4040
2.146	.1474	7.972	.4696
2.995	.2045	8.939	.5320
3.982	.2701	9.791	.5968
5.055	.3344		

M (25°)	κ 0°	κ 25°	κ 45°
0.0994	-	372	467
0.199	-	726	913
0.497	-	1723	2163
0.995	2187	3152	3911
0.501	2846	4287	5285
1.972	-	5044	6204
2.146	3581	5345	6561
2.504	3905	5752	7078
2.995	4183	6164	7631
3.982	4398	6573	8184
4.640	4381	-	-
4.963	-	6621	8341
5.055	4283	6582	8325
6.001	4195	6397	8152
7.028	3866	5999	7685
7.972	3468	5477	7112
8.939	2928	4796	6306
9.791	-	4215	5522

Slotte, 1883

%	d	t
	sat.sol	
18.66	1.125	10
33.56	1.2436	20
45.70	1.3598	30
58.94	1.5145	40

t	η 18.66%	η 33.56%	η 45.7%	η 58.94%
10	1475	1774	2289	3553
20	1165	1411	1803	2766
30	947	1152	1469	2224
40	787	962	1224	1835

Marignac, 1876

%	U	%	U
	21-53°		
3.18	0.9698	20.85	0.8251
6.17	0.9419	36.65	0.6964
11.63	0.8962		

Water + Ammonium chromate ($CrH_8N_2O_4$)

Gerasimov, 1934

%	f.t.	%	f.t.
20.01	0	34.15	50
25.35	20	41.80	75

Hill, Soth and Ricci, 1940

%	f.t.	%	f.t.
19.78	0	34.40	50
24.13	15	37.21	60
27.02	25	41.20	75
30.00	35		

Slotte, 1881

%	t	d
10.52	13	1.0633
19.75	13.7	.1197
28.04	19.6	.1727

t	η(water $^{10°}$=1000)		
	10.52%	19.75%	28.04%
10	1083	1203	1379
20	852	955	1101
30	-	788	912
40	578	661	769

Marignac, 1876

%	U	%	U
	21-52°		
4.06	0.9630	14.48	0.8767
7.27	0.9304	25.29	0.7967

Water + Acid ammonium chromate (CrH_5NO_4)

Slotte, 1881

t	η(water $^{10°}$=1000)		
	6.85%	13.0%	19.93%
10	989	991	1006
20	768	780	802
30	625	639	664
40	519	533	558

Water + Lithium chromate (Li_2CrO_4)

Tammann, 1885

%	p	%	p
	100°		
0	760.0	23.80	670.0
4.42	748.5	29.55	628.8
10.62	730.8	37.15	554.8
16.83	706.4		

Heydweiller, 1912

N	d	N	d
	18°		
0.0815	1.00319	0.823	1.04293
0.1629	1.00763	2.057	1.10593
0.4115	1.02111	4.115	1.20592

Heydweiller, 1909

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.35251
0.1	1.33529	2.0	1.37095
0.2	1.33727	4.0	1.40612
0.5	1.34307		

Rubien, 1911

N	n_D	N	n_D
	18°		
0	1.33327	0.823	1.34917
0.08145	1.33490	2.058	1.37196
0.1629	1.35695	4.115	1.40806
0.4115	1.34136		

Heydweiller, 1912

N	κ	N	κ
	18°		
0.0815	61.85	0.823	421.8
0.1629	113.4	2.057	763.1
0.4115	245.5	4.115	935.5

Water + Sodium chromate (Na_2CrO_4)

Tammann, 1885

%	p	%	p
	100°		
6.98	746.7	28.63	676.1
10.56	738.6	32.06	657.8
17.81	720.1	35.85	635.9
23.60	698.4	42.01	600.7
25.85	688.5		

Stokes, 1948

Isopiestic solutions at 25°

m_1	m_2	m_1	m_2
0.1127	0.1480	2.357	2.815
0.2158	0.2750	2.462	2.951
0.5533	0.6725	3.118	3.922
0.8564	1.016	3.489	4.475
1.070	1.249	3.866	5.150
1.428	1.653	4.363	6.105
2.121	2.490		

m_1 = m of Na_2CrO_4 m_2 = m of NaCl

Salkowski, 1901

%	f.t.	%	f.t.
43.65 6 aq.	17.7	45.27	23.2
44.12	19.2	45.75	24.7
44.64	21.2	46.28	26.6
46.47 4 aq.	28.9	47.08	31.2
46.54	29.7		

Jones, 1904 and Jones and Bassett, 1905

M	f.t.	M	f.t.
0.10	-0.45	0.50	-1.900
0.20	-0.85	0.60	-2.345
0.30	-1.23	0.80	-3.063
0.40	-1.604		

Benrath, 1942

%	f.t.	%	f.t.
56	140	64	280
57	165	66	296
58	190	67	305
59	210	68	315
60	225	70	335
61	245	73	350
62	260	76	372
63	270		

Slotte, 1881

%	t	d
5.76	17.4	1.0576
10.62	17.4	1.1125
14.81	20.7	1.1644

t	η 5.76 %	η 10.62 %	η 14.81 %
10	1527	1839	2271
20	1185	1413	1729
30	950	1130	1377
40	780	930	1121

Jones, 1904 and Jones and Bassett, 1905

M	d	M	d
	0°		
0.05	1.007500	0.50	1.069300
0.10	.015200	0.60	.083400
0.20	.028700	0.80	.110000
0.30	.042400	1.00	.136600
0.40	.056100		

Heydweiller, 1912

M	d	M	d
	18°		
0.1066	1.0063	1.066	1.0722
0.2132	1.0138	2.132	1.1426
0.533	1.0361	4.264	1.2740

Riss, 1939

%	d 30°	d 60°	d 90°
33.45	1.353	1.337	1.320
46.04	1.525	1.495	1.473
52.83	-	1.587	1.569
sat.sol.	-	-	1.320

%	η 30°	η 60°	η 90°
33.45	3230	1632	1060
46.05	8140	3420	2040
52.83	-	5940	3050
sat.sol.	-	-	1100

Morgan and Schramm, 1913

%	σ	%	σ
	30°		
0	71.03	39.96	84.57
4.45	71.64	43.47	87:86
7.70	72.18	45.25	89.73
12.67	73.35	46.78	91.26
19.35	74.96	47.72	92.68
24.05	76.48	49.81	95.36
29.96	78.74	53.12	97.46
34.91	81.52		

Jones, 1904 and Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.10	97.00	0.50	72.50
0.20	89.40	0.60	69.00
0.30	80.00	0.80	57.59
0.40	76.00	1.00	53.95

Heydweiller, 1912

M	κ	M	κ
	18°		
0.1066	87.2	1.066	607.2
0.2132	160.6	2.132	968.9
0.533	350.3	4.264	1238

Marignac, 1876

mol %	U
21-52°	
4	0.7810
2	0.8560
1	0.9134
0.5	0.9511

Water + Potassium chromate (K_2CrO_4)

Heterogeneous equilibria

Tammann, 1885

t	p				
	0%	20.22%	27.56%	32.84%	40.42%
42.99	64.8	61.7	59.7	57.1	-
46.37	77.9	73.4	70.8	69.1	-
49.70	91.2	86.4	83.4	81.2	-
57.32	131.9	124.5	120.4	117.2	-
59.20	144.0	137.2	132.2	128.7	-
62.18	165.2	157.2	152.0	147.6	-
64.96	187.3	177.5	171.6	166.5	-
68.08	215.1	204.7	197.5	191.9	-
69.75	231.3	219.6	211.6	206.3	-
72.40	259.2	246.6	239.1	231.3	-
74.66	285.2	270.9	261.9	255.3	-
76.81	311.9	296.2	286.8	279.5	265.3
82.50	393.0	372.7	360.8	350.9	332.9
84.37	423.2	402.1	389.5	377.8	358.3
87.26	473.5	-	437.8	434.6	406.1
87.87	484.9	459.8	446.0	434.8	412.9
89.50	516.1	490.9	465.0	461.9	439.8
91.57	558.0	530.3	513.5	499.9	474.8
93.53	600.5	572.2	555.3	546.2	514.1
97.50	645.7	614.6	596.1	580.2	550.5
99.79	754.3	715.5	695.2	675.5	642.4

%	p	%	p
	100°		
9.85	742.0	25.87	707.9
12.50	737.3	29.21	695.7
16.46	729.9	31.97	686.3
21.03	719.6	34.76	675.9
21.69	718.4		

Stokes, Wilson and Robinson, 1941

m	osmotic coefficient	activity coefficient
	25°	
0.1	0.805	0.456
0.2	.774	.382
0.3	.753	.340
0.4	.741	.313
0.5	.733	.292
0.6	.727	.276
0.7	.722	.263
0.8	.718	.253
0.9	.714	.243
1.0	.711	.235
1.2	.709	.223
1.4	.711	.214
1.6	.716	.207
1.8	.722	.201
2.0	.730	.196
2.5	.757	.190
3.0	.794	.190
3.5	.830	.191

Kremers, 1856

sat.sol. b.t.=107°

Baroni, 1893

%	b.t.	%	b.t.
1.24	100.110	12.06	100.794
2.55	100.152	14.72	101.007
4.06	100.248	18.53	101.331
6.19	100.306	23.23	101.818
8.79	100.562		

Rüdorff, 1862

%	f.t.
1.96	-0.4
3.85	-0.8
5.66	-1.2
7.41	-1.65

Nordenskjöld, 1869

%	f.t.	%	f.t.
38.08	0	42.85	63.6
38.31	10.0	44.35	93.6
39.87	27.37	44.99	106.1
41.28	42.1		

de Coppet, 1872

%	f.t.
9.09	-2.0
16.67	-3.9
23.08	-5.75
28.57	-7.8
33.33	-9.95

Rüdorff, 1872

%	f.t.	%	f.t.
9.09	-1.95	23.08	-5.70
16.67	-3.80	28.57	-7.80
19.35	-4.55	33.33	-10.00

Etard, 1894

%	f.t.	%	f.t.
39.7	34	42.6	96
40.3	53	44.0	120
41.8	79	45.4	157

Koppel and Blumenthal, 1907

%	f.t.	%	f.t.
4.33	-0.99	35.30	-11.37 E
5.77	-1.2	36.35	0
21.25	-4.3	39.37	30
23.86	-5.7	42.73	60
29.60	-7.12	47.03	105.8 b.t.
34.39	-10.35		

Yuskevich, Karzhavin and Shokin, 1926

%	t	%	t
24.1	0	48.00	40
44.30	10	50.7	50
45.15	20	53.6	60
46.45	40		

Benrath, Gjedebo and al., 1937

%	f.t.	%	f.t.
46.1	133	52.0	235
47.3	153	52.0	244
47.5	165	53.5	254
48.1	175	53.2	275
48.5	180	54.0	286
49.0	189	54.9	294
49.4	198	55.4	311
50.5	215	55.7	327

Properties of phases

Schiff, 1858 and 1859

%	d	%	d
	19.5°		
0	0.9983	21	1.1844
1	1.0063	22	.1944
2	.0144	23	.2046
3	.0226	24	.2149
4	.0308	25	.2253
5	.0391	26	.2358
6	.0475	27	.2464
7	.0558	28	.2571
8	.0645	29	.2679
9	.0732	30	.2787
10	.0819	31	.2899
11	.0907	32	.3013
12	.0996	33	.3129
13	.1095	34	.3246
14	.1176	35	.3364
15	.1268	36	.3482
16	.1361	37	.3602
17	.1455	38	.3723
18	.1551	39	.3845
19	.1641	40	.3968
20	.1745		

Fouqué, 1867

%	d	
	0°	16°
0.96	1.0077	1.0065
3.24	.0268	.0257
16.87	.1535	.0498

Slotte, 1881

24.26% d^{18}= 1.2335

Traube, 1885

c	d
	15°
0	0.9991
9.27	1.0795
17.35	.1530
30.95	.2923

Getman and Wilson, 1908

%	d	%	d
	20°		
0	0.9982	15	1.1287
3	1.0243	20	.1765
4	.0325	25	.2274
5	.0408	30	.2808
10	.0837	35	.3386

Heydweiller, 1909

N	d	N	d
	18°		
0.1	1.00642	1.0	1.0736
.2	.01422	2.0	.1452
.5	.03671	4.0	.2814

Schneider, 1910

M	d	M	d
	18°		
0	0.99862	0.50	1.07301
0.050	1.00635	0.99	.1441
0.099	.01407	1.98	.2791
0.198	.02896		

Rubien, 1911

M	d	M	d
	18°		
0.0507	1.00658	0.5019	1.07391
.1013	.01422	1.0039	.14533
.2526	.03708	1.9906	.28025

Flöttmann, 1928

%	t	d
38.49	15	1.3749
38.94	20	.3785
39.98	25	.3805

Yuskevich, Karzhavin and Shokin, 1926

%	t	d	%	t	d
24.1	0	1.237	48.00	40	1.539
44.30	10	.491	50.7	50	.579
45.15	20	.510	53.6	60	.612
46.45	30	.532			

Gibson, 1935

%	d	%	d
	25°		
0.000	0.99717	26.502	1.23661
13.782	1.11302	32.900	.30606
19.908	1.12048	37.689	.36223

Jones and Colvin, 1940

M	d	M	d
25°		0°	
0.0005	0.99716	0.0005	0.99995
.001	.99723	.001	1.00003
.002	.99739	.002	.00019
.005	.99786	.005	.00070
.01	.99863	.01	.00155
.02	1.00018	.02	.00315
.05	.00481	.05	.00799
.1	.01236	.1	.01593
.2	.02753	.2	.03181
.50178	.07199	.304656	.07814
1.0005	.14316	1.0078	.15149
1.9992	.27891	2.01707	.28991
2.52209	.34708	2.54422	.35890

Gibson, 1935

%	κ (1-1000 bars)
0.000	39.35
13.782	32.20
19.908	29.24
26.502	26.28
32.900	23.58
37.689	21.51

Slotte, 1881

t	η (water=1)
10	1.333
20	.065
30	0.879
40	.743

Schneider, 1910

M	η (water=1)	M	η(water=1)
	18°		
0.050	1.0082	0.50	1.0797
.099	.0142	0.99	.1820
.198	.0306	1.98	.4840

Jones and Colvin, 1940

M	η(water=1)	M	η(water=1)
25°		0°	
0.0005	1.00036	0.0005	1.00025
.001	.00053	.001	.00032
.002	.00090	.002	.00046
.005	.00170	.005	.00070
.01	.00284	.01	.00090
.02	.00492	.02	.00122
.05	.01062	.05	.00157
.1	.01940	.1	.00223
.2	.03732	.2	.00496
.30178	.09457	.304656	.02536
1.0005	1.20945	1.0078	.09270
1.9992	.53171	.948675	.32130
2.52209	.76931	2.01707	.33989
		.54422	.54108

Traube, 1885

c	σ
15°	
0	73.26
9.27	73.75
17.35	75.36
30.95	78.76

Getman and Wilson, 1908

%	n_D	%	n_D
	20°		
0	1.33298	15	1.36715
3	.33973	20	.37940
4	.34187	25	.39233
5	.34401	30	.40525
10	.35525	35	.41985

Rubien, 1911

M	n_D	M	n_D
	18°		
0.0507	1.33532	0.5019	1.35312
.1013	.33737	1.0039	.37192
.2526	.34345	1.9906	.40674

Heydweiller, 1913

N	n_D	N	n_D
	18°		
0	1.33327	1.0	1.35305
0.1	.33529	2.0	.37188
.2	.33732	4.0	.40709
.5	.34335		

Flöttmann, 1928

%	t	n_D (sat.sol.)
38.49	15	1.43267
38.94	20	.43276
39.98	25	.43288

Heydweiller, 1909

N	λ	N	λ
	18°		
0.1	100.5	1.0	79.5
0.2	94.4	2.0	72.0
0.5	86.5	4.0	59.9

Dewar and Fleming, 1897 and 1898

t	ε	t	ε
	36%		
-201.5	10.2	-135.7	14.0
-193.7	11.1	-127.0	14.6
-177.0	12.1	-122.0	15.3
-165.0	12.5	-112.4	16.9
-148.8	12.6	-100.0	22.3

Faasch, 1911

N	U	N	U
	18°		
0.489	0.939	1.994	0.810
0.990	0.898	3.986	0.703

Water + Magnesium chromate ($MgCrO_4$)

Hill, Glenn and al., 1940

%	f.t.	%	f.t.
32.06	0 (7+1)	34.94	10 (5+1)
33.87	10	35.20	15
34.78	15	35.39	25
		35.81	35
		36.82	50
		37.68	60
		39.96	75

Slotte, 1881

%	t	d
12.31	13.6	1.0886
21.86	14.5	.1641
27.71	13.6	.217

Heydweiller, 1912

N	d	N	d
	18°		
0.1234	1.00718	1.234	1.0803
.247	.01570	2.468	.1580
.617	.04034	4.936	.3034

Slotte, 1881

t	η(water^{10}=1) 12.31%	21.86%	27.71%
10	1.518	2.278	3.177
20	1.154	1.709	2.356
30	0.920	1.347	1.828
40	0.750	1.086	1.455

Heydweiller, 1912

N	ϰ	N	ϰ
	18°		
0.1234	67.7	1.234	421.7
.247	120.9	2.468	613.3
.617	253.6	4.936	603.4

Water + Rubidium chromate (Rb_2CrO_4)

Schreinemakers and Filippo, 1906

%	f.t.	%	f.t.
0.949	-0.60	36.65 E	- 7.
7.215	-1.10	38.27	0.
9.872	-1.57	40.225	+10.5
15.576	-2.40	42.422	20.
20.027	-3.25	44.114	30.
24.283	-4.44	46.13	40.
30.153	-5.55	47.44	50.
34.341	-6.71	48.90	60.4

Water + Bichromic acid ($H_2Cr_2O_7$)

Jones, 1904and Jones and Bassett, 1904 and 1905

M	f.t.	d	molecular conductivity
		0°	
0.025	-0.526	-	452.0
.05	-0.775	-	440.0
.10	-1.050	1.01270	436.0
.15	-1.610	.02050	428.0
.20	-2.22	.02760	420.0
.30	-2.89	.04150	410.0
.40	-6.78	.05550	388.9
.50	-11.47	.06920	374.3
1.00	-16.00	.13740	312.4
1.50	-22.50	.20380	256.2
2.00	-31.00	.27050	206.8
2.50	-42.00	.33520	157.3
3.00	-57.50	.39930	128.2
3.50	-	.46440	108.4
4.00	-	.52990	-

Water + Ammonium bichromate ($Cr_2H_8N_2O_7$)

Gerasimov, 1913

%	f.t.
15.37	0
26.23	20
42.03	50
52.13	75

Water + Novocaine bichromate ($Cr_2C_{26}H_{42}N_4O_{11}$)

Holleman and De Jong, 1940 (fig.)

%	L_1+L_2	$V+L_1+L_2$	%	L_1+L_2	$V+L_1+L_2$
0	0	-	60	93	74
1	63	-	70	81	74
9	74	74	74	74	74
20	92	74	80	81	-
30	98	74	90	105	-
40	99	74	100	133-135	-

Water + Lithium bichromate ($Li_2Cr_2O_7$)

Hartford and Lane, 1948

%	f.t.	%	f.t.
0	0.0	60.55	30.0
18.73	- 6.2	61.45	40.0
27.87	-12.0	62.57	50.0
32.38	-17.2	63.60	60.0
38.22	-24.0	64.62	70.0
44.07	-35.4	65.81	80.0
48.66	-49.8	67.20	90.0
53.50	-60	68.40	100.0
57.99	+ 0.8		

tr.t. : 63.02 % +42.1° (9+2) - (3+1)

Heydweiller, 1921

N	d	N	d
		18°	
0.2	1.01581	1	1.0780
0.5	1.03928	2	1.1570

Hartford and Lane, 1948

%	d	%	d
	30°		
0	0.9957	34.07	1.2575
3.75	1.0195	36.07	.2777
5.72	.0320	38.58	.3026
7.59	.0455	41.31	.3325
11.28	.0705	42.87	.3470
14.29	.0930	47.17	.3954
17.65	.1178	48.98	.4177
20.81	.1427	52.49	.4602
24.94	.1767	57.59	.5246
28.66	.2082	60.55	.567

Heydweiller, 1921

N	λ	N	λ
		18°	
0.2	73	1	58.7
0.5	66.1	2	47.7

Water (H_2O) + Sodium dichromate ($Na_2Cr_2O_7$)

Stanley, 1866

c	f.t.	c	f.t.
107.2	0	142.8	80
109.2	15	162.8	100
116.6	30	209.7	139

Jones, 1904 and Jones and Bassett, 1905

M	f.t.
0.10	-0.490
0.20	-0.946
0.30	-1.400
0.40	-1.872

Winslow and Hartford, 1941

%	f.t.		%	f.t.
0.00	0.0		66.06	32.2
22.05	-5.0		67.52	37.6
26.45	-5.4		67.98	40.6
37.47	-10.6		69.21	46.1
48.50	-19.7		71.38	57.2
57.32	-35.7		71.76	60.0
58.90	-39.0		72.15	60.9
59.96	-44.0		73.82	67.9
60.54	-47.1		73.76	68.3
60.77	-48.2	E	75.65	74.4
60.84	-45.4		78.41	72.2
61.73	-8.0		79.65	90.1
62.17	0.0		80.01	95.0
62.49	5.0			
65.01	25.0			

tr.t. 2aq. - 0 aq. : 79.18% 84.6°

Stanley, 1866

%	d	%	d
	at room t.		
1	1.007	26	1.175
2	.014	27	.185
3	.021	28	.193
4	.028	29	.201
5	.035	30	.208
6	.042	31	.216
7	.049	32	.224
8	.057	33	.231
9	.064	34	.238
10	.071	35	.245
11	.078	36	.252
12	.085	37	.259
13	.092	38	.266
14	.099	39	.273
15	.105	40	.280
16	.113	41	.287
17	.120	42	.280
18	.127	43	.287
19	.134	44	.294
20	.141	45	.300
21	.147	46	.307
22	.153	47	.313
23	.159	48	.319
24	.165	49	.325
25	.171	50	.343

Jones, 1904 and Jones and Bassett, 1905

M	d	M	d
	0°		
0.05	1.009500	0.60	1.112200
0.10	.019000	0.80	.149100
0.20	.037700	1.00	.184900
0.30	.056800	1.25	.228500
0.40	.075600		

Pesce, 1935

N	d	N	d
	25°		
6.410	1.56232	0.961	1.08660
4.372	.38637	0.918	.08262
2.048	.18498		

Jones, 1904 and Jones and Bassett, 1905

M	molecular conductivity	M	molecular conductivity
	0°		
0.05	100.8	0.30	85.7
0.10	95.0	0.40	81.5
0.20	89.0		

Water + Potassium dichromate ($K_2Cr_2O_7$)

Leopold and Johnston, 1927

m	t	p 0 %	p sat.sol.
0.482	23.66	21.92	21.43
.545	26.40	25.93	25.30
.698	32.82	37.35	36.05
.790	36.21	45.08	43.42
.920	40.85	57.88	55.48
1.210	50.77	96.10	91.55

Stokes, 1948

m	osmotic coefficient 25°
0.1	0.868
.2	.813
.3	.779
.4	.753
.5	.735

Baroni, 1893

%	b.t.	%	b.t.
1.62	100.074	7.27	100.336
3.43	100.158	9.25	100.436
5.48	100.255	10.79	100.517

Kremers, 1854

%	f.t.	%	f.t.
4.73	0	33.56	60
7.81	10	42.19	80
11.56	20	50.50	100
22.57	40		

Tilden and Shenstone, 1884

%	f.t.
56.20	117
60.59	129
66.74	148
72.43	180

Etard, 1894

%	f.t.	%	f.t.
4.1	-1	42.8	92
4.3	+1	44.0	97
5.6	6	48.0	104
6.1	7	52.0	120
7.2	12	54.4	130
8.5	15	60.8	150
10.4	20	62.8	157
14.2	29	66.6	178
16.6	36	76.9	215
28.2	57	89.7	291
30.2	61	91.8	312
32.0	65	97.4	360
39.9	70		

Koppel and Blumenthal, 1907

%	f.t.	%	f.t.
4.30	-0.63 E	31.24	60
4.43	0	51.96	104.8 b.t.
15.34	+30		

Le Blanc and Schmandt, 1911

%	f.t.
5.52	4.81
15.17	30.10
17.77	35.33

Yuskevich, Karzhavin and Shokin, 1926

%	f.t.	%	f.t.
61.30	0	67.65	40
62.70	10	70.10	50
64.50	20	71.70	60
66.20	30		

Rakovski and Babajeva, 1931

%	f.t.	%	f.t.
4.47	0	31.30	60
10.97	20	42.20	80
20.83	40	50.00	98

Kremers, 1855

%	d	%	d
19.5°			
0	0.998	9	1.063
1	1.005	10	.071
2	.013	11	.078
3	.020	12	.093
4	.028	13	.085
5	.035	14	.100
6	.041	15	.108
7	.048		
8	.054		

Kasankin, 1891

%	d	%	d
16-17°		32-34°	
0	0.9989	0	0.9947
sat.sol.	1.0652	sat.sol.	1.0189

%	h*	%	h
16-17°		32.34°	
0	26.44	0	25.92
sat.sol.	25.51	sat.sol.	24.09

*h= capillary rise

Heydweiller, 1921

N	d
18°	
0.1	1.01010
.2	.02012
.5	.05005

Yuskevich, Karzhavin and Shokin, 1926

%	t	d	%	t	d
61.30	0	1.689	67.65	40	1.805
62.70	10	.721	70.10	50	.838
64.50	20	.756	71.70	60	.882
66.20	30	.778			

(sat.sol.)

Flöttmann, 1928

%	t	d
8.893	15	1.0635
10.822	20	.0768
12.98	25	.0916 (sat.sol.)

Flöttmann, 1928

%	t	n_D (sat.sol.)
8.893	15	1.35028
10.822	20	.35345
12.98	25	.35685

Heydweiller, 1921

N	λ
18°	
0.1	98.2
.2	93.1
.5	85.4

Water + Rubidium bichromate ($Rb_2Cr_2O_7$)

Stortenbeker, 1914

% I	% II	f.t.
5.42	4.96	18
6.94	6.55	24
9.08	8.70	30
13.22	12.90	40
18.94	18.77	50
28.1	27.3	65

Water + Calcium bichromate ($CaCr_2O_7$)

Hartford, Lane and Meyer, Jr., 1950

%	f.t.	%	f.t.
0.00	0	50.29	-12.2 (9+2)
8.00	-1.2	53.01	0.4
18.55	-3.1	57.13	20.0
26.25	-6.1	59.56	30.0
33.69	-11.0	61.00	35.0
40.59	-17.2	62.47	40.0
46.74	-26.2	63.40	45.0 (3+1)
47.50	-27.1	64.21	50.0
48.90	-29.9	65.74	60.0

E: 49.3% -31.3°

Tarasenkov and Konopkina, 1954

%	f.t.	%	f.t.
5.09	-0.4	47.76	-27.6
9.75	-1.5	47.80	-27.8
18.94	-3.1	50.80	-35.4
28.46	-7.6	52.50	-40.0 E ice + (y+1)
38.40	-15.1		
53.47	0 (6+1)	64.35	50 (4+1)
54.60	8.1	65.54	60
55.34	10	66.37	70 (x+1)
57.87	20 (5+1)	68.32	80
59.92	30	70.60	90
61.14	35	72.70	100
62.60	40		

E : 52.5 %

Hartford, Lane and Meyer Jr., 1950

%	d	%	d
	30°		
0.00	0.9957	34.15	1.3348
4.53	1.0330	39.70	.4074
10.07	.0798	48.90	.5469
19.07	.1646	54.58	.6469
27.76	.2568		

Tarasenkov and Konopkina, 1954

t	d	t	d
-20	1.6599	40	1.610
0	.6416	60	.5944
+20	.6257	75	.5832

Water + Dimethyl molybdate ($C_2H_6MoO_4$)

Rosenheim and Bertheim, 1903

%	f.t.	%	f.t.
2.575	-0.510	11.635	-2.220
5.016	-1.022	12.090	-2.665
6.781	-1.430	14.096	-3.085
8.820	-1.900		

Water + Ammonium molybdate ($H_8N_2MoO_4$)

Rakshit, 1935

%	d	%	d
	20°		
1	1.00423	5	1.03287
2	.01141	10	.06731
3	.01804	30	.20188

Water + Potassium Octacyanomolybdate ($K_4C_8MoN_8$)

Brubaker, 1956

m	osmotic coeff.	m	osmotic coeff.
	25°		
0.05105	0.6337	0.8093	0.4831
.06562	.6157	.8338	.4826
.2031	.5524	.8988	.4824
.2504	.5447	.9299	.4804
.2615	.5354	.9609	.4841
.3539	.5225	.9989	.4853
.4258	.5183	1.048	.4862
.5733	.4900	.135	.4899
.6485	.4867	.246	.4902
.7093	.4843	.429	.5094
.7591	.4835		

Water + Sodium molybdate (Na_2MoO_4)

Tammann, 1885

%	p	%	p
	100°		
13.11	735.6	34.48	663.5
21.26	716.0	40.57	625.7
29.65	686.8	44.68	593.1

Funk, 1900

%	f.t.	%	f.t.
	10 aq.		
30.63	0	35.58	6
33.83	4	38.16	9
	2 aq.		
39.28	10	41.27	51.1
39.27	15.5	45.57	100
39.82	32		

sat. sol. 18° d = 1.437

Traube, 1895

%	d
15°	
0	0.99911
6.31	1.0555
11.84	1.1085
19.91	1.1928
22.18	1.2194

Water + Potassium molybdate (K_2MoO_4)

Tammann, 1885

%	p
100°	
16.40	735.4
27.56	706.6
36.15	674.7
47.01	610.1
52.63	562.9

Water + Magnesium molybdate ($MgMoO_4$)

Ricci and Linke, 1951

%	f.t.	d	
3.94	-0.45	-	
6.28	-0.75	-	
9.38	-1.25	-	
11.55	-1.67 E	1.111	
12.94	5	.127	(7+1)
13.50	7	.134	
14.31	10	.143	
14.61	11	.145	
15.15	12.7 E	.151	
14.43	5	1.143	(5+1)
14.85	10	.147	
15.26	15	.151	
15.59	20	.150	
15.89	25	.159	
16.29	30	.161	
16.88	36	.164	
17.62	43	.160	?
17.74	45	.178	
18.40	50	.180	
19.00	55	.185	
19.48	58	.184	
19.55	59	.192	
19.90	62	.195	
19.85	60.8 E	.192	
9.38	95	1.072	(2+1)
10.79	90	.084	
12.09	85	.096	
13.70	80	.121	
15.30	75	.141	
16.13	72	.147	
16.51	71	.149	
18.83	64	.177	
19.12	63	.182	
19.96	60	.196	

Water + Ammonium luteophosphomolybdate ($H_{24}Mo_{18}N_6O_{62}P_2$)

Nikitina and Sokolova, 1955

%	f.t.	%	f.t.
34.46	0	50.74	25
37.88	5	53.64	30
45.06	15	55.95	35
48.90	20	58.93	40

%	d	η (water=1)
	20°	
19.93	1.1776	1.2255
30.89	.3064	.4136
41.03	.4482	.7048
46.26	.4491	.7025

Water + Sodium tungstate (Na_2WO_4)

Tammann, 1885

%	p	%	p
	100°		
8.39	751.4	42.64	665.4
16.29	739.2	45.46	650.9
26.27	716.9	46.75	644.5
34.02	697.7	49.11	630.1
36.97	688.4		

Funk, 1900

%	f.t.	%	f.t.
	10 aq.		
30.60	-5	36.54	0
31.87	-4	39.20	+3
32.98	-3.5	41.02	+5
34.52	-2		
	2 aq.		
41.67	- 3.5	43.98	43.5
41.73	+ 0.5	47.65	80.5
42.27	+21	49.31	100

sat.sol. 18° d = 1.573

Franz, 1872

%	d	%	d
	24.5°		
0	0.997	20.49	1.192
0.89	1.001	21.38	.201
1.78	.009	22.28	.212
2.67	.018	23.17	.224
3.56	.026	24.06	.236
4.45	.033	24.95	.247
5.38	.042	25.84	.258
6.27	.049	26.73	.270
7.17	.056	27.62	.285
8.06	.065	28.55	.301
8.91	.072	29.40	.317
9.80	.081	30.29	.331
10.69	.089	31.18	.345
11.58	.098	32.07	.360
12.47	.107	32.96	.377
13.36	.116	33.85	.393
14.29	.127	34.74	.410
15.18	.136	35.64	.426
16.08	.144	36.53	.441
16.97	.153	37.42	.456
17.82	.163	38.31	.472
18.71	.173	39.20	.488
19.60	.182		

Traube, 1895

%	d
15°	
0	0.9991
4.14	1.0373
10.10	1.0954
19.35	1.2012
33.18	1.3944

Water + Sodium metatungstate ($Na_2W_4O_{13}$)

Tammann, 1885

%	p
100°	
19.36	753.6
30.40	750.0
47.34	741.0
68.87	696.3
76.86	627.8

Water + Potassium tungstate (K_2WO_4)

Tammann, 1885

%	p	%	p
	100°		
16.47	742.4	44.98	663.1
26.01	724.9	49.87	635.4
34.06	708.0	56.25	588.7
36.76	694.8	61.01	544.6

Water + Tungsten anhydride (WO_3)

Scheibler, 1861

%	d	%	d
	17.5°		
0	0.9987	26	1.3006
1	1.0078	27	.3160
2	.0170	28	.3317
3	.0204	29	.3478
4	.0359	30	.3642
5	.0456	31	.3810
6	.0554	32	.3982
7	.0654	33	.4158
8	.0756	34	.4338
9	.0860	35	.4521
10	.0966	36	.4710
11	.1075	37	.4903
12	.1185	38	.5100
13	.1297	39	.5301
14	.1412	40	.5507
15	.1529	41	.5718
16	.1650	42	.5933
17	.1771	43	.6154
18	.1896	44	.6379
19	.2025	45	.6608
20	.2156	46	.6844
21	.2290	47	.7085
22	.2428	48	.7330
23	.2567	49	.7580
24	.2710	50	.7837
25	.2856		

Gerlach, 1859

%	d	%	d
	17.5°		
0	0.9927	30	1.3642
5	1.0456	35	.4521
10	.0966	40	.5507
15	.1529	45	.6609
20	.2156	50	.7838
25	.2856		

Water + Metatungstic acid ($W_4O_{13}H_2W$)

Sobolev, 1896

%	f.t.	%	f.t.
0.29	-0.018	9.53	- 0.465
0.59	-0.034	17.58	- 0.879
1.19	-0.062	25.02	0
2.39	-0.117	40.09	+22
4.76	-0.228	45.06	+43.5

%	t	d
25.02	0	1.6025
40.09	22	2.5239
45.06	43.5	3.6503

Water + Phosphotungstic acid ($W_{12}O_{40}H_3$)

Sobolev, 1896

c	b.t.	c	b.t.
13.702	100.080	19.926	100.109
13.702	100.079	25.192	100.129

c	f.t.	c	f.t.
0.2983	-0.0146	5.930	-0.1360
.5090	-0.0230	6.6227	-0.1536
.7188	-0.0317	7.954	-0.1750
2.0040	-0.0570	9.916	-0.2110
3.9930	-0.0960	13.702	-0.2870

%	t	d
12.24	0	1.1890
28.97	22	.6913
30.65	43	.8264
40.78	92	2.5813

Water + Sodium phosphotungstate ($Na_3PO_4.12WO_3$)

Sobolev, 1896

c	f.t.
22.04	0
59.65	22
98.184	93

c=g $Na_3PO_4.12WO_3.21$ aq in 100 cc water

Water + Rhenic acid (ReO_4H)

Feit, 1931

%	d	%	d
	17°		
0	1.000	43.23	1.55
2.93	.025	45.94	.60
5.71	.050	48.18	.65
8.37	.075	50.29	.70
10.91	.100	52.57	.75
15.65	.150	54.44	.80
20.00	.200	56.22	.85
24.40	.250	57.90	.90
28.46	.300	59.59	.95
31.85	.350	61.00	2.00
35.00	.400	62.44	.05
37.93	.450	63.81	.10
40.67	.500	65.12	.15

Q . METALLIC SALTS + OTHER SOLVANTS .

LV . METALLIC SALTS + ORGANIC SOLVANTS .

Butane (C_4H_{10}) + Aluminium bromide ($AlBr_3$)

Heldman and Thurmond, 1944

mol %	f.t.	mol %	f.t.
9.5	28.3	56.9	67.8
11.5	32.2	60.2	71.1
16.1	38.8	63.0	72.7
16.3	39.2	66.2	74.1
19.4	43.2	72.4	77.2
22.7	46.7	74.6	79.8
24.4	48.6	77.9	80.9
25.1	48.7	77.1	81.4
28.1	51.4	77.9	82.5
30.9	53.4	86.2	84.3
37.6	57.6	86.4	86.3
48.6	64.1	90.7	89.9
53.7	67.0	100	97.5

Pentane (C_5H_{12}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol %	sat.t.
58.0	142.0
51.5	157.1
50.5	161.8
38.7	182.0
33.5	190.5
25.9	199.8

Isopentane (C_5H_{12}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol %	sat.t.
61.9	133.6
62.1	136.1
54.2	166.9
52.8	170.8
51.2	174.4
24.8	215.0

Hexane (C_6H_{14}) + Aluminium bromide ($AlBr_3$)

Boedeker and Oblad, 1947

wt %	mol %	f.t.	wt %	mol %	f.t.
31.00	12.7	30.6	70.56	43.7	60.3
34.55	14.6	33.1	70.73	43.9	60.5
41.98	19.0	40.5	81.68	59.1	69.8
46.47	21.9	44.1	82.47	60.4	70.3
51.46	25.4	48.5	82.18	59.9	70.7
53.36	26.6	49.9	84.81	64.4	72.6
55.11	28.4	50.0	92.91	81.3	82.4
58.80	31.6	53.9	95.00	86.1	86.6
62.10	34.6	55.0	100.00	100.0	97.5

Hexane (C_6H_{14}) + Stannic iodide (SnI_4)

Dice and Hildebrand, 1928

mol %	sat.t.	mol %	sat.t.
25.31	138.1	45.26	148.7
28.87	143.4	53.44	145.4
32.25	146.4	59.16	138.2
37.14	148.8	42 mol %	C.S.T.=149.4

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
30.5	143.7	31.8	141.1
34.3	146.7	38.3	146.2
38.2	147.5	39.1	146.4
40.7	147.6	42.0	146.4
42.9	147.6	46.2	146.0
43.5	147.6		
47.3	147.0		
54.2	142.4		

2-Methylpentane (C_6H_{14}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
29.8	179.9	43.3	185.1
34.7	184.8	47.7	183.4
38.8	185.4	56.0	173.7

3-Methylpentane (C_6H_{14}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
29.6	143.8	41.7	147.8
34.5	146.7	46.4	147.0
39.2	147.6	50.0	145.7

2-3 Dimethylbutane (C_6H_{14}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
35.2	166.3	42.2	166.7
39.3	166.9	47.3	165.3
39.7	166.8	49.8	163.2

Heptane (C_7H_{16}) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Vold, Leggett and Mc Bain, 1940 (fig.)

%	clear.p.	%	clear.p.
2	220	80	269
10	244	89	270
40	260	100	297

%	f.t.	%	f.t.
0.3	30	1.0	200
0.4	50	20.0	250
0.5	90	86.4	280
0.8	150		

Heptane (C_7H_{16}) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol %	clear.p.
32.2	14.3	240

Heptane (C_7H_{16}) + Stannic iodide (SnI_4)

Dice and Hildebrand, 1928

mol %	sat.t.	mol %	sat.t.
28.36	127.6	51.42	136.4
32.20	130.9	52.09	136.3
34.34	131.6	56.06	135.4
35.95	134.1	57.00	135.9
37.20	134.8	58.47	134.8
43.40	136.4	59.45	134.7
45.73	136.7	61.05	131.5
47.45	136.3	62.40	132.0
47.76	136.9	65.48	128.9
48.61	136.7	68.18	127.2

48 mol % C.S.T. = 136.8

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
33.8	131.9	48.4	135.7
34.2	132.6	48.8	135.4
34.8	132.4	49.0	136.0
35.0	131.7	49.5	135.4
41.2	135.0	54.8	132.5
43.3	135.5	56.4	134.6
44.2	135.4	58.0	133.4
45.9	135.9	60.2	130.0
46.8	135.5		

Heptane (C_7H_{16}) + Titanium chloride ($TiCl_4$)

Neff and Hickman, 1955 (fig.)

mol %	p	
	50°	60°
0	140	210
25	120	180
50	95	145
75	70	105
100	45	65

2-Methylhexane (C_7H_{16}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol%%	sat.t.	mol %	sat.t.
34.5	156.9	48.9	160.3
41.8	160.0	52.2	159.8
44.7	160.3	61.4	153.4

3-Methylhexane (C_7H_{16}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
34.5	141.0	48.8	144.0
41.5	143.9	51.7	143.8
44.6	144.1	60.1	140.1

Stannic tetrachloride ($SnCl_4$) (b.t. = 101.65) + Hydrocarbons

Lecat, 1949

2nd Comp.			Az		
Name	Formula	b.t.	%	b.t.	Dt mix or sat.t.
Octane	C_8H_{18}	125.75	80	113.2	- 0.7 (10%)
Diisobutyl	C_8H_{18}	-	60	107.5	- 1.8 (50%)
Toluene	C_7H_8	110.75	46.5	109.15	- 1.2 (67%)

Octane (C_8H_{18}) + Stannic chloride ($SnCl_4$)

Campbell and Hickman, 1953

mol % L	mol % V	b.t.	mol % L	mol % V	b.t.
75.7	81.5	114	24.2	33.2	120
36.2	45.6	118	5.0	7.9	124

Octane (C_8H_{18}) + Stannic iodide (SnI_4)

Dice and Hildebrand, 1928

mol %	sat.t.	mol %	sat.t.
36.02	127.3	54.18	132.3
38.24	129.4	58.32	131.5
44.54	131.1	62.97	130.2
45.41	131.2	64.30	129.4

52 mol % C.S.T. = 132.0

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
35.4	127.2	57.6	131.5
44.4	131.1	60.2	130.7
51.7	131.6	48.5	131.5

Isooctane (C_8H_{18}) + Stannic iodide (SnI_4)

Dice and Hildebrand, 1928

mol %	sat.t.	mol %	sat.t.
30.29	183.4	53.80	194.9
31.46	185.4	59.20	193.6
36.92	191.6	62.53	188.1
45.92	194.9	62.56	191.6
47.58	194.9	66.04	187.5
51.93	195.4	69.22	181.5

50 mol % C.S.T. = 195.3

3-Methylheptane (C_8H_{18}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
37.6	135.3	51,7	138.3
44.7	137.9	55.7	137.8
47.1	138.1	66.0	133.3
51.0	138.3		

2,2,4-Trimethylpentane (C_8H_{18}) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol %	clear.p.
47.4	26.6	248
73.3	44.3	258

2,2,4-Trimethylpentane (C_8H_{18}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol %	sat.t.	mol %	sat.t.
34.2	185.0	48.8	195.1
41.2	194.1	53.0	195.0
44.3	194.8	63.0	189.3

Hexadecane ($C_{16}H_{34}$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggett, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
0.1	30	1.4	200
0.1	50	86.4	250
0.3	90	94.5	280
0.5	150		

Hexadecane($C_{16}H_{34}$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Stross and Abrams, 1951

%	tr.t_1	Q trans	tr.t_2	Q trans
3.0	95-100 109-111 121-126	- - -	150-152	- - -
9.3	92-98 104-111 122-126	6960	153-156	1170
26.4	92-100 104-112	7150	151-155	1380
34.9	94-99 106-112 124-126	7020	152-155	1420
50.0	93-99 110 125-126	7500	154-156	960
63.6	94-100 109-111 126	7650	151-156	950
76.6	86-96 107-108 125-126	7200	150-159	1345
86	88-94 108-110 128	6600	178-180	1470
100	86-96 112-118 129-130	7000	198	1170

%	tr.t_3	Q trans	tr.t_4	Q trans	tr.t_5	Q tr.
3.0	210	-	237-241	-	-	-
9.3	212-216	820	236-241	1770	-	-
26.4	210-217	470	235-241	1590	-	-
34.9	210-216 234-241	2630	-	-	-	-
50.0	218	250	234-240 257	1600	-	-
63.6	217-220	450	241-245	1890	-	-
76.6	226-229	520	248-250	1890	-	-
86	241-248 252.255	2460	-	-	-	-
100	243-251	1920	271-273	20	280	140

Stross and Abrams, 1950

%	clear.p.	%	clear.p.
2	167	50	248
9	215	67	259
18	220	71	255
30	240	100	282

Hexadecane($C_{16}H_{34}$)+ Calcium stearate ($CaC_{36}H_{70}O_4$)

Vold and Smith, 1951

X-ray diffraction of the monohydrate

Dotriacontane ($C_{32}H_{66}$) + Stannic chloride ($SnCl_4$)

Hildebrand and Wachter, 1949

%	f.t.	%	f.t.
98.70	30.1	91.63	40.8
96.13	37.8	72.3	53.8

Nujol (paraffin oil) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Vold, Leggett and Mc Bain, 1940 and 1942

Clearing point curve.

Propylene (C_3H_6) + Silver nitrate ($AgNO_3$)

Francis, 1951

mol%	f.t.	mol%	f.t.
44.4	40	43.5	36
43.5	36	43.0	19.4
42.5	25	42.4	0
42.6	23.3	42.3	-3.5
42.2	19.4	42.4	-3.5
40.7	0	40.5	-14
40.3	0	38.3	-26
40.2	0	42.2	-13
39.6	-10	42.4	-13
38.8	-19	42.4	-13.5
38.3	-25	40.5	-23
		38.7	-26
		38.3	-26

Diisobutylene (C_8H_{16}) + Stannic iodide (SnI_4)

Bocharov and Obolenzev, 1941

mol%	sat.t.
23.8	117.5
39.5	128.3
50.8	129.0
58.5	127.4

Cyclohexene (C_6H_{10}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	%	f.t.
19.1	-25	64.5	35
20.4	-15	74.0	45
30.0	-5	83.6	55
37.1	+5	92.8	65
45.1	15	100.0	73
54.3	25		

Cyclohexene (C_6H_{10}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	%	f.t.
11.7	-5	55.1	65
12.8	5	64.5	70
15.1	15	76.2	75
19.0	25	84.4	80
24.1	35	90.7	85
31.3	45	95.8	90
41.0	55	100.0	94

Cyclohexane (C_6H_{12}) + Sodium stereate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%	clear.p.	%	mol%	clear.p.
0.44	0.2	115	39.6	16.2	243
7.9	2.5	160	45.4	19.8	above310
11.5	3.8	195	58.2	29.2	" 310
18.3	6.2	222	72.0	43.3	" 300
25.1	9.1	232	83.0	59.2	" 307
33.2	12.8	240	93.3	80.5	269
			100.0	100.0	285

Cyclohexane (C_6H_{12}) + Aluminium bromide ($AlBr_3$)

Leighton and Wilkes, 1948

%	f.t.	%	f.t.
25.1	6.2	57.9	38.7
27.6	8.8	59.4	39.8
35.2	17.2	64.2	44.5
45.1	26.4	76.0	57.0
47.4	28.5	79.5	60.3
55.0	36.0	80.4	61.7
56.4	37.6	100.0	97.5

Cyclohexane (C_6H_{12}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

mol%	f.t.	mol% L_1	mol% L_2	sat.t.
0.00	6.4	5.9	92.3	70
0.08	6.0	8.3	90.2	80
0.4	20	11.3	87.1	90
0.9	30	15.1	82.5	100
1.6	40	20.1	76.1	110
2.6	50	28.3	65.4	120
3.8	60	34.7	54.7	124
		44.1	44.1	125 C.S.T.

Cyclohexane (C_6H_{12}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	sat.t.
0	6.4	-	-
0.2	6.4	-	-
0.2	6	-	-
1.2	20	6	-
2.5	30	6	-
4.2	40	6	-
6.5	50	-	-
9.7	60	-	-
13.7	70	-	-
19.5	"	-	80
25.6	"	-	90
32.3	"	-	100
40.4	"	-	110
51.1	"	-	120
58.9	"	-	124
68.0	"	-	125.5
76.7	"	-	124
83.2	"	-	120
89.6	"	-	110
92.7	"	-	100
94.7	"	-	90
96.1	"	-	80
97.0	"	-	-
100.0	73	-	-

Cyclohexane (C_6H_{12}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	sat.t.
0	6.4	-	-
0.3	6	-	-
1.4	20	6	-
2.4	40	6	-
3.7	50	6	-
5.3	50	6	-
7.1	60	-	-
9.5	70	-	-
12.5	80	-	-
16.2	90	-	-
17.4	92.5	-	100
21.2	"	-	110
25.8	"	-	120
31.0	"	-	130
36.4	"	-	140
42.0	"	-	150
47.8	"	-	160
53.8	"	-	170
62.3	"	-	173
66.3	"	-	175
74.0	"	-	173
82.9	"	-	170
86.3	"	-	160
90.9	"	-	-
92.7	"	-	-
94.0	"	-	-
95.0	"	-	-
95.8	"	-	-
96.5	"	-	-
97.1	"	-	-
97.6	"	-	-
100.0	94	-	

Cyclohexane (C_6H_{12}) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1946

mol%	sat.t.
40.5	110.2
47.4	112.4
55.8	115.1
56.6	115.4

Cyclohexane (C_6H_{12}) + Tin tetraisopropyl ($C_{12}H_{28}Sn$)

Schulze, 1951

mol%	d	mol%	d
	25°		
0	0.77390	61.363	1.06793
15.668	.88447	70.082	.09082
27.347	.94589	89.897	.13358
36.839	.98681	100	.15169
45.557	1.01915		

Methylcyclohexane (C_7H_{14}) + Stannic chloride ($SnCl_4$)

Lecat, 1949

%	b.t.	Dt. min.
0	101.15	-
10	-	-0.6
15	100.8 Az	-
100	113.85	-

Methylcyclohexane (C_7H_{14}) + Stannic iodide (SnI_4)

Bocharov and Obolenzev, 1946

mol%	sat.t.
40.9	114.7
51.7	116.7
61.2	119.7

Dimethyl-1,3-cyclohexane (C_8H_{16}) + Stannic chloride ($SnCl_4$)

Lecat, 1949

%	b.t.
0	120.7
80	112.5 Az
100	113.85

Benzene (C_6H_6) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%	clear.p.
20.1	6.5	222
29.9	10.5	above 238

Benzene (C_6H_6) + Silver perchlorate ($AgClO_4$)

Hill, 1922

%	f.t.	d	
		15°	
100	480	-	
65.6	160	-	
64.6	159	-	
63.0	145	-	tr.t.
62.6	140	-	
60.0	138.5	-	(1+1)
50.5	115.5	-	
40.1	92	-	
32.2	80.3	1.164	sic
10.07	50	0.895	
5.00	25	0.906	
3.44	5.12	0.909	
0	5.48	-	

Benzene (C_6H_6) + Nickel ricinoleate ($NiC_{36}H_{66}O_4$)

Walden, 1914

%	D f.t.
10.41	-0.117
11.38	-0.133
14.04	-0.203
17.06	-0.298

Benzene (C_6H_6) + Aluminum chloride ($AlCl_3$)

Eley and King, 1951

%	f.t.	%	f.t.
4.88	108.6	40.50	151.8
4.83	108.7	41.64	152.4
5.84	111.3	47.83	155.3
13.66	127.5	53.07	158.4
16.52	132.0	57.18	160.9
19.86	136.0	60.06	164.4
22.31	138.5	100.00	192.0
30.03	145.1		

Benzene (C_6H_6) + Aluminium bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	5.7	77.4	50
10.0	4.5	83.0	60
20.0	3.0	87.5	70
27.4	1.8	91.2	80
35.3	10	95.3	90
46.5	20	97.5	94
59.0	30	100	96
70.0	40		

Plotnikov and Gratseanskii, 1939

mol%	f.t.	E	tr.t.	
0	5.48	-	-	
6.26	3.90	3.78	-	
10.05	4.20	3.78	-	
11.53	5.55	3.8	-	
14.19	8.98	3.8	-	
15.99	11.68	3.8	-	
17.86	15.62	3.8	-	
24.80	22.98	3.8	-	
25.55	24.38	3.8	-	
28.76	27.38	3.8	-	
45.4	38.98	-	37.25	(1+1)
50.4	45.38	-	37.00	"
50.7	45.48	-	"	"
52.99	46.88	-	"	"
62.2	57.38	-	"	"

Eley and King, 1951

%	f.t.	%	f.t.
55.54	28.2	87.78	68.4
63.59	34.3	89.91	72.0
65.07	35.8	96.63	91.6
65.49	35.9	96.67	92.7
68.57	36.6	97.21	93.6
72.78	40.1	97.92	95.4
80.11	51.5	98.51	96.9
85.33	60.8	99.04	97.3
85.63	60.9	100.00	98.1

Andreevskii, 1955

%	mol%	f.t.	E
0	0	5.52	-
2.0	0.58	5.31	-
2.7	0.81	5.23	-
3.47	0.9	5.22	-
3.5	1.05	5.16	-
5.5	1.67	4.955	-
6.77	1.95	4.87	-
7.0	2.16	4.78	-
8.03	2.35	4.62	-
10.0	3.0	4.42	-
10.44	3.3	4.37	-
12.75	4.0	4.02	1.92
13.75	4.4	3.92	-
15.3	5.0	3.72	1.92
17.0	5.7	3.52	-
18.8	6.0	3.25	1.92
20.33	6.95	2.55	-
20.43	7.0	2.5	-
22.0	7.6	2.82	-
22.93	7.9	2.62	-
24.2	8.54	2.52	-
24.2	8.54	2.41	1.96
24.6	8.7	2.42	1.92
25.3	9.0	2.27	-
26.54	9.54	2.07	-
27.4	9.94	1.98	1.96
27.49	9.95	1.92	1.92
28.6	10.5	3.55	-
31.65	12.0	4.95	-
43.77	18.5	17.12	1.92
44.21	18.8	15.22	1.92
54.92	26.4	-	1.92

Plotnikov, Sheka and Yankelevich, 1939

%	d	ε
	20°	
0.00	0.8790	2.282
0.05	0.8792	2.287
0.63	0.8820	2.309
1.81	0.8880	2.336
2.15	0.8902	2.338
5.34	0.9110	2.372
7.06	0.9217	2.375
8.21	0.9300	2.372
9.28	0.9300	2.375
20.16	0.9373	2.480
24.57	1.0225	2.497
37.10	1.0595	2.556
40.89	1.1817	2.608
43.66	1.2237	2.634
47.89	1.2577	2.700
50.69	1.3070	2.750

Benzene (C_6H_6) + Aluminum bromide-Lithium chloride complexe ($LiAl_2ClBr_6$)

Gorenbein, 1956

mol%	D f.t.	mol%	Df.t.
3.85	-1.58	1.57	-0.84
2.93	1.35	1.51	0.82
2.39	1.13	1.31	0.78
1.85	0.95	1.28	9.74
1.67	0.88		

Benzene (C_6H_6) + Aluminum bromide-Sodium chloride complexe ($NaAl_2ClBr_6$)

Gorenbein, 1956

mol%	D f.t.	mol%	Df.t.
2.93	-0.50	1.33	-0.44
2.43	0.48	1.17	0.42
2.07	0.46		

Benzene (C_6H_6) + Aluminum bromide-Potassium chloride complexe (KAl_2ClBr_6)

Gorenbein, 1956

mol%	D f.t.	mol%	D f.t.
3.24	-0.39	0.87	-0.26
2.53	0.37	0.77	0.24
2.02	0.35	0.71	0.23
1.66	0.33		

Benzene (C_6H_6) + Aluminum bromide.Hydrogene sulfide ($AlBr_3.H_2S$)

Jakubsohn, 1925

%	D.f.t.	%	D.f.t.
1.1	-0.202	3.7	-0.621
1.9	.336	5.7	.949
2.4	.413	7.0	1.154
2.5	.432	12.2	.891
3.6	.628	22.53	4.153

%	κ	%	κ
	25°		
20.17	0.0081	38.19	0.6321
23.17	.0227	41.85	.7863
24.34	.0338	45.20	1.129
29.44	.1072	47.89	2.185
32.45	.1656	49.48	.427
35.67	.3917	53.99	3.540

Benzene (C_6H_6) + Aluminum iodide (AlI_3)

Eley and King, 1951

%	f.t.	%	f.t.
7.38	29.5	55.17	106.6
8.91	34.3	59.57	110.7
21.12	61.6	62.66	112.9
27.79	73.4	70.62	125.5
36.34	80.4	84.07	144.5
38.90	98.0	86.50	149.0
43.85	92.6	100.00	188.8
57.10	105.6		

Benzene (C_6H_6) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.	E
0	5.6	-	76.8	75	1
7.3	4	1	80.5	77.5	1
19.4	1	-	85.3	79	-
24.6	10	1 (1+2)	89.5	77.5	62
30.5	20	1	91.5	75	62
37.0	30	1	93.5	70	62
44.1	40	1	96.0	62	-
52.0	50	1	97.9	67.5	62
60.6	60	1	100	73	-
70.7	70	1			

Kurnakov, Krotkov and Oksman, 1915 and Kivinakov, 1924

mol%	d	η	mol%	d	η
		75°			
0	0.8178	337	66.6	2.0824	1855
25	1.2911	606	75	.2293	2051
33.4	.4330	786	85	.4173	2257
50	.8137	1301	100	.6724	2359

Traube, 1895

%	d
20°	
7.333	0.93209
10.134	.95059
13.160	.97548

Benzene (C_6H_6) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E
0	5.6	-
8.3	4.5	-
12.5	15	4.5
17.2	25	"
23.0	35	"
30.5	45	"
39.0	55	"
48.8	65	"
60.5	75	"
67.2	80	"
74.3	85	"
83.0	90	"
86.5	91.5	"
90.2	92.5	-
92.8	91.5	85
93.8	90	85
96.3	85	-
98.0	90	85
98.9	92	85
100	94	-

Benzene (C_6H_6) + Antimony tribromide. Aluminium bromide ($SbBr_3.AlBr_3$)

Gorenbein, 1940

%	d			
	10°	20°	30°	40°
80.03	-	-	2.210	2.192
73.92	-	1.9954	1.9777	1.9599
73.02	-	.995	.978	.959
71.68	-	.922	.904	.885
68.60	-	.824	.808	.793
67.35	1.807	.791	.777	.762
64.577	.7328	.7175	.7014	.6844
63.76	.706	.694	.678	.664
58.669	.5930	.5780	.5623	.5467
58.48	.588	.573	.558	.544
52.73	.4734	.4586	.4449	.4297
52.05	.460	.446	.431	.417
48.94	.405	.391	.377	.363
45.93	.367	.354	.340	.327
40.47	.2747	.2620	.2480	.2347
38.38	.245	.233	.219	.205
36.22	.218	.205	.192	.180
30.21	.148	.135	.123	.112
23.36	.075	.062	.051	.040
15.6	.002	0.992	0.983	0.968
10.84	0.964	.953	.943	.931
0	.885	.875	.865	.855

Gorenbein, 1940

%	η 10°	20°	30°	40°
80.025	-	-	13061.0	8648.7
73.92	-	8811.1	5561.2	4155.0
71.68	-	6914.7	4603.1	3103.0
68.60	-	4701.5	3212.4	2381.9
67.35	6133.0	4088.9	2992.9	2192.8
64.58	5024.4	3662.3	2744.2	2109.4
58.669	3201.2	2430.4	1900.5	1520.5
52.73	2290.2	1794.0	1450.1	1190.1
52.05	2118.6	1630.7	1334.3	1083.0
45.93	1688.5	1348.1	1112.9	920.6
40.47	1414.2	1160.7	970.9	823.0
38.38	1352.0	1123.0	949.5	791.9
36.32	1289.8	1077.0	888.3	755.3
30.2	1108.9	926.2	766.1	657.9
23.36	1009.0	850.2	715.1	621.4
15.6	882.9	751.5	655.4	569.0
10.84	854.0	729.7	620.5	539.7
0	799.5	652.8	555.8	486.6

%	κ 10°	20°	30°	40°
80.025	-	-		15.77
78.40	5.36	8.43	12.06	16.27
76.96	5.97	8.94	12.49	16.50
75.50	-	-	-	16.28
74.20	6.14	9.02	-	-
72.0	6.48	8.96	-	-
71.20	6.54	8.89	-	-
70.93	6.52	8.96	-	-
68.40	6.31	8.22	-	-
67.90	6.52	8.58	10.80	-
67.50	6.43	8.46	10.59	12.81
67.35	6.25	8.04	-	-
64.58	-	-	8.839	10.42
59.30	4.80	5.83	6.695	7.57
58.669	4.108	4.875	-	-
56.3	3.86	4.52	5.101	5.59
52.73	2.641	2.959	3.218	-
52.05	2.90	-	-	-
51.10	2.38	2.56	2.70	-
47.10	1.67	1.76	1.783	1.79
41.73	0.715	0.710	0.619	0.562
40.47	-	0.4746	0.442	0.4055
38.38	-	0.430	0.39	-
36.32	-	-	0.133	-
33.20	0.166	0.154	-	0.104
30.9	-	-	-	0.032

Benzene (C_6H_6) + Stannic chloride ($SnCl_4$)

Schulze and Hock, 1914

%	P kg. 125°	120°	p mm 30°
0	3.35	1.82	116
10	3.27	.82	"
20	3.17	.79	"
30	3.03	.74	114
40	2.87	.67	110
50	2.68	.57	104
60	2.46	.46	94
70	2.23	.29	84
80	1.96	1.19	71
90	1.69	0.920	56
100	1.38	0.688	37

%	d	%	d
	20°		
0	0.877	60	1.358
10	.919	70	.496
20	.978	80	.660
30	1.050	90	.883
40	.139	100	2.2184
50	.240		

Trifonov, 1924

%	d	%	d
	26°		
0	0.8702	82.50	1.7756
24.44	1.0191	100.00	2.2062
54.43	1.2954		

Kurnakov, Perelmuter and Konov, 1915-16, and Kurnakov, 1924

mol%	d 25°	70°
0	0.8729	0.8241
25	1.2819	1.2014
50	1.6226	1.4832
75	1.9038	1.8060
100	2.2132	2.0962

mol%	η 25°	70°
0	608	363
25	600	375
50	638	399
75	713	479
100	919	600

Anosov, 1926

%	n_D	%	n_D
	20°		
0.00	1.50209	82.50	1.50578
24.44	.50030	100.00	.51228
54.43	.50145		

Trifonov, 1924

%	χ	%	χ
	26°		
0	-0.802	82.50	-0.544
24.44	-0.733	100.0	-0.484
54.43	-0.635		

Benzene (C_6H_6) + Stannic bromide ($SnBr_4$)

Kapustinskii and Kessler, 1956 (fig.)

mol%	f.t.	mol%	f.t.
0	5.58	50	+2
10	0	60	6
20	-5	70	13
30	-9	75	15.3 tr.t.
34.1	-9.9 E	90	24
40	-4	100	29.40

Benzene (C_6H_6) + Stannic iodide (SnI_4)

Bocharov and Obolentsev, 1947

mol%	sat.t.	mol%	sat.t.
36.8	99.6	51.4	109.2
39.1	102.2	65.2	115.8

Dorfmen and Hildebrand, 1927

%	mol%	f.t.
9.76	1.330	10.00
15.17	2.181	25.00
23.02	3.593	40.00

Benzene (C_6H_6) + Tin tetraisopropyl ($SnC_{12}H_{28}$)

Schulze, 1951

mol%	d	mol%	d
	25°		
0	0.87343	52.686	1.07892
9.627	.93328	87.028	.13615
30.211	1.01949	87.813	.13706
35.177	.03473	100	.15172

Benzene (C_6H_6) + Iron pentacarbonyl $Fe(CO)_5$

mol%	d	mol%	d
20°			
0.00	0.8790	18.35	1.0125
2.01	.8950	40.48	.1580
4.07	.9100	100.00	.4606
8.35	.9401		
20°			
0.00	0.8790	11.15	0.9650
1.33	.8900	29.42	1.0856
2.69	.9002	100.00	1.4606
5.47	.9220		

mol%	n_F	n_D	n_C
20°			
0.00	1.5129	1.50118	1.49638
1.33	.5121	.50014	.49535
2.69	.5115	.49938	.49463
5.47	.5107	.49835	.49335
11.15	.5098	.49541	.48984
29.42	.5098	.49541	.48984
100.00	.5380	.51745	.50970

mol %	ε
20°	
0.00	2.292
2.01	2.305
4.07	2.317
8.35	2.331
18.35	2.368
40.48	2.437
100.00	2.602

Toluene (C_7H_8) + Sodium stearate ($C_{18}H_{35}O_2Na$)

Smith and Mc Bain, 1947

%	mol%	clear.p.	%	mol%	clear.p.
0.93	0.3	133	55.9	29.0	287
9.1	3.14	135	60.3	33.0	331
18.9	6.5	225	77.9	53.3	335
36.0	15.5	244	89.8	74.1	257
45.4	21.2	235	94.6	85.0	264

Toluene (C_7H_8) + Silver perchlorate ($AgClO_4$)

Hill and Miller, 1925

%	f.t.	$d_{(f.t.)}$	%	f.t.	$d_{(f.t.)}$
54.60	75	1.665	42.89	16.5	1.388
52.68	50	.576	42.00	16	.375
50.30	25	.523	26.41	0	.129
tr.t.	22.6	-	6.01	-24.1	0.920
44.11	18	.417	0	-73.5	.854

Toluene (C_7H_8) + Magnesium oleate ($MgC_{36}H_{66}O_4$)

Nelson and Pink, 1954 (fig.)

c	η	c	η
25°			
0	540	15	1130
5	650	17	1300
10	820		

Toluene (C_7H_8) + Calcium oleate ($CaC_{36}H_{66}O_4$)

Nelson and Pink, 1954 (fig.)

c	η
25°	
0	540
5	760
10	1210
15	2160

Toluene (C_7H_8) + Zinc laurate ($ZnC_{24}H_{46}O_4$)

Martin and Pink, 1948 (fig.)

%	sat.t.	C.S.T.	%	sat.t.
0.1	75	-	1.0	88
0.2	80	90.8	3.5	90
0.4	85	-	14.0	91

Toluene (C_7H_8) + Zinc decoate ($ZnC_{20}H_{38}O_4$)

Martin and Pink, 1948

C.S.T. = 86.0°

Toluene (C_7H_8) + Zinc myristate ($ZnC_{28}H_{54}O_4$)

Martin and Pink, 1948

C.S.T. = 93.0°

Toluene (C_7H_8) + Zinc stearate ($ZnC_{36}H_{70}O_4$)

Martin and Pink, 1948 (fig.)

%	sat.t.	C.S.T.	%	sat.t.
0.1	75	-	4.0	95
0.2	80	-	5.5	96
0.5	85	-	16.0	97
1.5	90	97.0		

Toluene (C_7H_8) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
16.4	-15	76.5	50
18.3	-10	82.2	60
23.7	0	87.2	70
32.1	10	91.4	80
42.5	20	95.7	90
56.0	30	97.8	94
68.8	40	100.0	96

Toluene (C_7H_8) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	tr.t.	
0	-93	-	-	
1.1	-94	-	-	
2.0	-80	-94	-	(1+1)
3.1	-70	-94	-	
4.8	-60	-94	-	
7.3	-50	-94	-	
11.0	-40	-94	-	
15.8	-30	-	-	
21.5	-20	-	-	
29.7	-10	-	-	
41.5	0	-	-	
40.8	+6	-	-	
57.8	11	-	-	
47.8	-8	-	-	(1+2)
52.5	+1	-	-	
57.8	11	-	-	
62.8	20	-	11	
69.0	30	-	"	
73.2	35	-	"	
78.0	40	-	"	
79.8	41.5	-	"	
83.1	42.5	-	-	
84.8	41.5	40	-	
85.8	40	-	-	
89.0	50	40	-	
92.7	60	40	-	
95.0	65	40	-	
97.8	70	40	-	
98.8	71.5	-	-	
100.0	73	-	-	

Toluene (C_7H_8) + Antimony tribromide.Aluminum bromide ($SbBr_3.AlBr_3$)

Gorenbein, 1940

%	d 0°	d 10°	d 20°	d 30°	d 40°
75.41	2.076	2.060	2.045	2.029	2.012
73.89	2.022	2.006	1.990	1.975	1.959
68.67	1.855	1.841	1.826	1.811	1.796
65.51	1.768	1.753	1.740	1.725	1.710
62.84	1.698	1.683	1.669	1.656	1.642
54.43	.506	.494	.481	.468	.455
48.68	.406	.394	.382	.368	.355
45.15	.349	.337	.325	.312	.300
38.75	.254	.243	.232	.220	.209
30.99	.157	.147	.136	.125	.114
28.68	.130	.120	.110	.099	.089
23.83	.081	.069	.060	.049	.031
17.35	.019	.009	0.999	0.989	0.979
8.49	0.945	0.936	.923	.914	.908
4.01	-	-	-	.883	.877
0	.884	.876	.867	.858	.844

%	η 0°	η 10°	η 20°	η 30°	η 40°
75.41	35622	19910	12052	-	-
73.89	27189	16500	10848	7540	5429
68.67	12284	8205.6	5876	4368	3325
65.51	8426	6009	4428	3394	2631
62.84	6601	4740	3587	2811	2298
54.43	3378	2592	2078	1708	1427
48.68	2514	1973	1610	1360	1151
45.15	2113.1	1708.5	1412	1196	1020
38.75	1662	1359.5	1162	1029.7	862.3
30.99	1342	1117.4	954	840	727
28.68	1222	1035	890	779	683
23.83	1106	929	805	708	619
17.35	1024	882	770	684	605
8.49	885	-	652	574	511
4.01	-	-	-	548	491
0	775	670	591	528	469

%	κ 0°	κ 10°	κ 20°	κ 30°	κ 40°
73.89	4.696	7.030	10.017	13.499	17.402
68.67	6.453	8.877	11.720	14.811	18.293
65.51	7.453	9.723	12.334	-	17.925
62.84	7.119	9.142	11.266	13.620	15.917
54.43	5.908	7.186	8.370	9.585	10.708
48.68	4.621	5.449	6.263	6.985	7.646
45.15	3.472	4.033	4.539	4.974	5.290
38.75	2.033	2.322	2.559	2.746	2.832
30.99	0.818	-	1.095	1.173	1.227
28.68	-	0.682	0.735	-	-
23.83	-	-	-	0.407	0.394
17.35	-	-	-	.121	.114
8.49	-	-	-	.005	.0035

Toluene (C_7H_8) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	tr.t.	
0	-93	-	-	
1	-93.5	-	-	
2.4	-80	-93.5	-	(1+1)
4.0	-70	-93.5	-	
6.2	-60	-93.5	-	
8.9	-50	-	-	
12.4	-40	-	-	
17.5	-30	-	-	
25.7	-20	-	-	
37.5	-10	-	-	
45.3	-5	-	-	
53.1	-1	-	-	
61.3	10	-	-1	(1+2)
69.4	20	-	-1	
78.0	30	-	-	
80.6	40	-	30	
83.5	50	-	30	
86.6	60	-	30	
90.0	70	-	-	
93.8	80	-	-	
95.7	85	-	-	
98.0	90	-	-	
100	94	-	-	

Toluene (C_7H_8) + Stannic chloride ($SnCl_4$)

Lecat, 1949

%	b.t.	Dt. mix.
0	110.75	-
33	-	-1.2
53.5	109.15	Az
100	113.85	-

De Carli, 1931

%	d 20°	η 20°	η 30°
100.00	2.2295	1045	812
90.19	1.9208	990	785
82.15	.7170	968	841
61.54	.3790	768	680
51.44	.2481	721	-
41.75	.1621	750	710
30.70	.0812	825	-
23.08	0.9929	801	683
11.60	.9272	782	-
0.00	.8660	584	573

Toluene (C_7H_8) + Stannic iodide (SnI_4)

Dorfman and Hildebrand, 1927

%	mol%	f.t.
9.99	1.605	10.00
14.39	2.507	25.00
21.93	3.967	40.00

Ethylbenzene (C_8H_{10}) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%		clear.p.
1.36	0.5		140
10.3	4.1		195
19.0	8.1		227
27.8	12.6		236
37.4	18.3		240
46.8	24.8		250
66.1	42.2	above	291
78.9	58.4		258
88.2	73.6		258
95.0	87.5		260

Ethylbenzene (C_8H_{10}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	tr.t.	
0.3	-93	-	-	(1+1)
0.6	-70	-93	-	
1.1	-50	-93	-	
2.5	-30	-93	-	
4.1	-20	-93	-	
7.0	-10	-	-	
11.4	0	-	-	
18.8	+10	-	-	
23.9	15	-	-	
30.3	20	-	-	
36.9	25	-	-	
44.4	30	-	-	
55.0	35	-	-	
60.4	37	-	-	
68.1	39	-	-	
74.1	37	-	35	
77.4	35	-	-	
80.4	33	-	-	
83.0	40	33	-	
87.2	50	33	-	
91.8	60	33	-	
94.6	65	33	-	
98	70	33	-	
100	73	-	-	
56.3	15	-	-	(2+1)
65.7	25	-	-	
70.2	30	-	-	
77.4	35	-	-	
78.4	36	-	35	
81.0	37	-	35	
81.8	36.8	-	-	

Ethylbenzene (C_8H_{10}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	tr.t.	
0.4	-93	-	-	
1.0	-70	-93	-	
1.5	-60	-93	-	
2.2	-50	-93	-	
3.2	-40	-93	-	
4.8	-30	-93	-	
7.5	-20	-93	-	
12.0	-10	-	-	
18.7	0	-	-	
29.2	+10	-	-	
37.0	15	-	-	
46.3	20	-	-	
57.5	25	-	-	(1+1)
69.7	29	-	-	
74.0	40	-	29	
78.2	50	-	29	
82.7	60	-	29	
87.3	70	-	29	
92.1	80	-	29	
94.7	85	-	29	
97.7	90	-	29	
100.0	94	-	-	

Propylbenzene (C_9H_{12}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	
1.3	-80	-	(1+1)
2.0	-70	-	
3.7	-60	-	
6.0	-50	-	
9.4	-40	-	
14.0	-30	-	
22.5	-20	-	
29.4	-15	-	
38.4	-10	-	
46.0	-20	-	
47.9	-10	-	
49.0	-5	-	
50.3	0	-5	
53.3	+10	"	
57.1	20	"	
60.0	-1.5	"	
61.4	30	"	
66.2	40	"	
71.5	50	-	
77.2	60	-	
83.3	70	-	
89.8	80	-	
96.7	90	-	
100.0	94	-	

Propylbenzene (C_9H_{12}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	tr.t.	
2.2	-67	-	
8.4	-32	-	
	-42		
11.0	-36	-	
14.1	-24	-	
18.4	-18	-	
23.6	-22	-	
28.1	-9	-	
	-18		
0.6	-70	-	(1+2)
1.4	-60	-	
2.4	-50	-	
5.5	-40	-	
10.1	-30	-	
17.2	-20	-	
21.5	-15	-	
26.6	-10	-	
32.7	-5	-	
40.4	0	-	
51.6	+5	-	
57.5	7	-	
68.2	8.5	-	
1.5	-70	-	(1+1)
2.7	-60	-	
5.5	-50	-	
9.6	-40	-	
16.0	-30	-	
25.7	-20	-	
30.5	-15	-	
38.6	-10	-	
48.6	-5	-	
59.2	0	-	
65.3	+1.5	-	anh.
66.3	1.0	1	
68.6	10	1	
71.4	20	1	
74.5	30	1	
78.5	40	1	
80.7	45	1	
83.3	50	1	
86.1	55	1	
89.1	60	1	
92.5	65	1	
97.0	70	1	
100.0	73	-	

Cumene (C_9H_{12}) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%	clear.p.	%	mol%	clear.p.
11.3	5.1	222	33.4	17.5	238
23.1	11.2	228	43.7	34.8	245

Butylbenzene ($C_{10}H_{14}$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%	clear. p.
9.2	4.5	-
19.9	10.5	193-4
31.1	17.6	226
38.8	23.2	239

Isoamylbenzene ($C_{11}H_{16}$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	%	f.t.	%	f.t.
4.0	-80 (1+1)	24.2	-45 (1+2)	54.9	-21
7.2	-70	31.0	-35	56.0	-10
11.7	-60	38.7	-25	57.4	0
17.5	-50	47.2	-15	63.3	10
25.4	-40	56.8	-5	67.5	30
30.4	-35	62.8	0	72.6	40
36.3	-30	70.0	5	79.0	50
44.4	-25	75.5	7.5	87.1	60
51.6	-22			91.7	65
60.5	-20.5			97.3	70
64.6	-22			100	73

Isoamylbenzene ($C_{11}H_{16}$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	%	f.t.
4.5	-70 (1+1)	32.5	-17
6.1	-60	33.5	-10
8.3	-50	35.6	0
11.5	-40	38.2	+10
16.6	-30	41.6	20
21.0	-25	45.8	30
27.0	-20	51.3	40
32.5	-17	57.5	50
37.3	-15	65.0	60
44.9	-13	73.8	70
		78.8	75
		84.0	80
		89.4	85
		95.3	90
		100.0	94

o-Xylene (C_8H_{10}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	-29	-	-	
9.0	-32	-35	-	
14.0	-35	-	-	
17.5	-30	-35	-	(1+1)
24.8	-20	"	-	
33.4	-10	"	-	
43.4	0	"	-	
55.0	+10	"	-	
68.1	19.5	-	-	
61.5	8	-	-	
68.1	19.5	-	-	
71.3	25	-	19.5	(1+2)
75.7	30	-	"	
78.6	32.5	-	"	
81.0	33.5	-	-	
81.9	32.5	31.5	-	
82.5	31.5	-	-	
84.8	40	31.5	-	
88.0	50	"	-	
90.2	55	"	-	
92.4	60	"	-	
94.8	65	"	-	
96.5	68	"	-	
98.5	71	-	-	
100.0	73	-	-	

o-Xylene (C_8H_{10}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	-29	-	-	
10.5	-33	-	-	
17.0	-20	-33	-	(1+1)
24.6	-10	"	-	
34.5	0	"	-	
41.3	5	"	-	
48.7	10	"	-	
56.7	15	"	-	
65.8	20	"	-	
77.2	24	-	-	
78.6	22.5	-	-	
74.5	0	-	-	
76.2	10	-	-	
78.0	20	-	-	
80.0	30	-	22.5	
82.2	40	-	"	
84.7	50	-	"	
87.3	60	-	"	
90.1	70	-	"	
93.5	80	-	"	
97.7	90	-	-	
100.0	94	-	-	

o-Xylene (C_8H_{10}) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%	clear.p.
10.2	4.1	200
19.9	8.5	216
27.6	12.5	229
38.1	18.7	240

m-Xylene (C_8H_{10}) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%	clear.p.
10.4	4.2	-
20.0	8.6	223
29.4	13.5	230
41.1	20.7	244

m-Xylene (C_8H_{10}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	
0	-57	-	
7.5	-60.5	-	
10.3	-55	-60.5	(1+1)
15.8	-45	"	
22.0	-35	"	
29.0	-25	"	
36.9	-15	"	
46.2	-5	"	
61.2	+5	-	
67.1	+7.5	-	
44.5	-15	-	(1+2)
48.5	-5	-	
53.1	+5	1.5	tr.t.
58.7	15	"	
65.7	25	"	
70.3	30	"	
73.8	33	"	
77.7	36	"	
81.0	38	-	
83.7	36.5	-	
79.4	19	-	
82.0	30	-	
83.7	36.5	-	
84.8	40	36.5	tr.t.
87.7	50	"	
89.4	55	"	
91.5	60	"	
94.0	65	"	
95.8	68	"	
97.2	70	"	
99.1	72	-	
100.0	73	-	

m-Xylene (C_8H_{10}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	(1+1)
0	-57	-	-
5.5	-59.2	-	-
10.0	-45	-59.2	
14.2	-35	"	-
20.0	-25	"	-
27.0	-15	"	-
38.8	-5	"	-
46.7	0	"	-
56.6	5	-	-
68.2	10	-	-
77.2	13.5	-	-
72.2	-5	-	-
73.8	+5	-	-
75.6	15	12.5 tr.t.	-
77.6	25	"	-
79.8	35	"	-
82.3	45	"	-
85	55	"	-
87.9	65	"	-
91.2	75	"	-
95.3	85	"	-
97.9	90	-	-
100.0	94	-	-

m-Xylene (C_8H_{10}) + Stannic chloride ($SnCl_4$)

De Carli, 1931

%	d	η	
	20°	20°	30°
100	2.2295	1045	812
80.54	1.6890	915	710
50.98	.2794	943	636
41.75	.1810	793	614
32.03	.0590	767	590
0.00	0.8622	627	549

m-Xylene (C_8H_{10}) + Stannic iodide (SnI_4)

Dorfman and Hildebrand, 1927

%	mol%	f.t.
8.62	1.572	10.00
13.42	2.558	25.00
20.00	4.062	40.00

p-Xylene (C_8H_{10}) + Sodium stereate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%	clear.p.
19.5	8.4	220
28.6	13.1	232
43.7	22.5	244
56.3	32.6	216
61.9	37.8	above 280
87.8	73.0	261
90.8	76.6	263

p-Xylene (C_8H_{10}) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	14	79.6	60
11.4	12.5	85.6	70
25.0	10.2	90.9	80
35.7	20	95.4	90
47.2	30	97.7	94
61.2	40	100.0	96
72.2	50		

p-Xylene (C_8H_{10}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
11.7	11.7	-	-	
17.5	20	-	11.7	
26.0	30	-	"	(1+1)
31.3	35	-	"	
37.3	40	-	"	
43.8	45	-	"	
52.3	50	-	"	
62.7	55	-	"	
68.1	56	-	-	
0	14	-	-	
11.7	11.7	7	-	
20.7	10	7	-	
32.8	7	-	-	
37.8	15	7	-	
44.0	25	"	-	
50.3	35	"	-	
56.5	45	"	-	
62.7	55	"	-	(1+2)
66.1	60	-	55	
70.8	65	-	"	
74.6	67.5	-	"	
81.0	70	-	-	
85.6	67.5	58	-	
88.1	65	"	-	
90.3	61.5	"	-	
92.0	58	-	-	
93.3	61.5	58	-	
95.0	65	-	-	
97.2	69	58	-	
100.0	73	-	-	

p-Xylene (C_8H_{10}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	14	-	-	
16.6	12	10	-	
28.0	10	-	-	
36.0	20	10	-	(1+2)
44.6	30	"	-	
53.8	40	"	-	
63.5	50	"	-	
74.0	60	"	-	
81.0	65	"	-	
83.8	66.5	"	-	
87.3	67.5	-	-	
88.3	66.5	-	-	
89.5	70	-	66.5	
91.4	75	-	"	
93.4	80	-	"	
95.7	85	-	"	
97.0	88	-	"	
98.4	91	-	"	
100.0	94	-	-	

Martin and Pink, 1948

Xylene (C_8H_{10}) + Zinc salts

2nd comp.	C.S.T.
Zinc decoate ($ZnC_{20}H_{38}O_4$)	89.0°
Zinc laurate ($ZnC_{28}H_{46}O_4$)	93.7°
Zinc myristate ($ZnC_{28}H_{54}O_4$)	96.0°
Zinc stearate ($ZnC_{36}H_{70}O_4$)	98.5°

Xylene (C_8H_{10}) + Aluminum bromide ($AlBr_3$)

Plotnikov and Dibrova, 1940

%	κ
80.0	1.03
83.3	0.74
86.2	0.46

Plotnikov and Gitman, 1940

after h hours	κ	after h hours	κ
	31.9 mol%	18°	
0	2.54	192	7.24
24	4.04	288	8.69
48	4.70	312	8.88
72	5.24	336	8.92
96	5.77	360	8.99
144	6.53	384	9.07
168	6.98		
	29.2 mol%	25°	
1	0.68	190	10.58
22	3.90	195	10.48
27	4.98	214	10.58
93	10.61	238	10.86
141	10.61	267	11.26
147	10.61	357	11.31
165	10.65	362	11.55
171	10.52		
	29.0 mol%	30°	
0.5	0.5	124	9.14
4.2	1.71	144	9.17
46	6.81	148	9.21
52	7.97	168	9.24
95	8.81	172	9.24
100	8.88	193	9.25
120	9.03		

Cymene ($C_{10}H_{14}$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	-75	-	-	
2.0	-76.5	-	-	
4.8	-60	-76.5	-	
7.5	-50	-76.5	-	
10.8	-40	"	-	
15.0	-30	"	-	
21.0	-20	"	-	(1+1)
30.0	-10	"	-	
36.8	-5	-	-	
41.0	-3.5	-		(1+2)
46.1	+10	-	-3.5	
51.8	20	-	-3.5	
60.0	30	-	-3.5	
65.7	35	-	-	
76.4	40	-	-	
63.7	0	-	-	
66.5	10	-	-	
69.4	20	-	-	
72.5	30	-	-	
76.4	40	-	-	
81.2	50	-	40	
87.0	60	-	40	
90.8	65	-	40	
95.7	70	-	40	
100.0	73	-	-	

Cymene ($C_{10}H_{14}$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	-75	-	-	
2.0	-77	-	-	
2.6	-70	-77	-	(1+1)
4.2	-60	"	-	
6.1	-50	"	-	
8.7	-40	"	-	
12.3	-30	"	-	
17.8	-20	"	-	
27.0	-10	"	-	
42.3	0	-	-	
51.5	+5	-	-	
72.86	10	-	-	
49.1	-7	-	-	
51.5	+5	-	-	
56.0	20	-	5	
59.8	30	-	5	
64.1	40	-	5	
69.2	50	-	5	
75.0	60	-	5	
81.4	70	-	5	
88.4	80	-	-	
96.4	90	-	-	
100.0	94	-	-	

Cymene ($C_{10}H_{14}$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Smith and Mc Bain, 1947

%	mol%	clear.p.
9.5	4.7	118
19.5	10.3	144
27.6	15.2	171
39.3	23.4	-

Mesitylene (C_9H_{12}) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	-54.4	-	-	
1.5	-55.6	-	-	
3.0	-40	-55.6	-	(1+1)
4.8	-30	"	-	
7.0	-20	"	-	
9.7	-10	"	-	
14.2	0	"	-	
20.3	+10	"	-	
29.0	20	-	-	
39.7	30	-	-	
46.0	35	-	-	
51.4	38	-	-	
54.2	45	-	-	
59.0	55	-	38	(1+2)
65.4	65	-	"	
70.0	70	-	"	
73.1	73	-	"	
79.2	75.5	-	-	
83.7	73	-	58.5	
87.0	70	-	"	
89.7	65	-	"	
92.4	58.5	-	-	
94.0	63	-	58.5	
96.7	68	-	"	
98.0	70	-	"	
100	73	-	-	

Mesitylene (C_9H_{12}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	-54.4	-	-	
2.1	-55.2	-	-	
2.7	-40	-55.2	-	(1+1)
3.6	-30	"	-	
5.7	-20	"	-	
9.0	-10	"	-	
16.0	0	-	-	
25.4	+10	-	-	
35.5	20	-	-	
46.5	29	-	-	
54.2	40	-	29	(2+1)
61.7	50	-	"	
65.8	55	-	"	
70.2	60	-	"	
75.8	65	-	-	
78.7	67	-	-	
82.5	69	-	-	
85.8	69.5	-	-	
87.7	69	-	-	
84.5	61	-	-	
88.3	70	-	-	
92.7	80	-	-	
98.0	90	-	-	
100	94	-	-	

1,2,4 Trimethylbenzene (C_9H_{12}) + Antimony trichloride ($SbCl_3$)

Menshutkin,

%	f.t.	E	tr.t.	
0	-57.4	-	-	
11.0	-58.5	-60	-	
18.6	-60	-	-	
23.6	-45	-60	-	(1+1)
27.8	-35	"	-	
33.3	-25	"	-	
40.2	-15	"	-	
45.0	-10	"	-	
50.7	-5	-	-	
57.5	-1	-	-	
65.3	0	-	-	
48.4	-15	-	-	
50.7	-5	-	-5	(1+2)
53.1	+5	-	"	
55.8	15	-	"	
58.7	25	-	-	
62.2	35	-	-	
66.8	45	-	-	
69.7	50	-	-	
75.3	55	-	-	
79.2	56	-	-	
82.5	55	51	-	
87.5	51	-	-	
89.2	55	51	-	
91.2	60	"	-	
93.9	65	"	-	
97.5	70	"	-	
100.0	73	-	-	

1,2,4 Trimethylbenzene (C_9H_{12}) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	-57.2	-	-	
9.7	-58.8	-	-	
11.0	-50	-58.8	-	(1+1)
13.2	-40	"	-	
16.2	-30	"	-	
21.0	-20	"	-	
31.0	-10	"	-	
38.3	-5	"	-	
47.6	0	-	-	
58.4	5	-	-	
63.5	7	-	-	
75	10	-	-	
55.5	-15	-	-	
58.7	-5	-	-	
62.6	+5	-	-	(1+2)
67.4	15	-	7	
73.0	25	-	7	
79.1	33	-	7	
85.7	36	-	-	
73.5	-10	-	-	
74.5	0	-	-	
75.7	+10	-	-	
77.1	20	-	-	
78.7	30	-	-	
80.6	40	-	-	
82.8	50	-	-	
85.4	60	-	-	
88.4	70	-	-	
92.4	80	-	-	
97.4	90	-	-	
100	94	-	-	

1,2,4 Trimethylbenzene (C_9H_{12}) + Stannic chloride ($SnCl_4$)

De Carli, 1931

%	d	η	
	20°	20°	30°
-	2.2295	1040	810
18.60	1.7231	961	812
29.13	.5281	933	789
38.84	.3787	886	768
49.12	.2638	873	758
66.04	.1041	861	745
79.73	.0011	878	747
89.29	0.9401	888	766
100.00	.8786	1011	812

Diphenyle ($C_{12}H_{10}$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	
0	70.5	-	
14.0	65	50	
25.3	60	"	
33.4	55	"	
40.0	50	-	
45.2	55	50	(1+2)
51.4	60	"	
59.0	65	"	
70.7	70	"	
74.6	71	-	
78.3	70	57	
85.5	65	"	
88.9	57	-	
93.1	65	57	
97.0	70	"	
100.0	73	-	

Diphenyle ($C_{12}H_{10}$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	tr.t.	
0	70.5	-	-	
20.6	65	47	-	
35.7	60	"	-	
47.2	55	"	-	
54.3	50	"	-	
57.4	47	-	-	
60.8	50	47	-	(1+2)
68.5	55	"	-	
75.3	57.5	"	-	
82.7	60.5	-	-	
76.6	50	-	-	
81.6	60	-	-	
86.5	70	-	60.5	
91.5	80	-	"	
97.3	90	-	"	
100.0	94	-	-	

Diphenyle ($C_{12}H_{10}$) + Antimony triiodide (SbI_3)

Menshutkin, 1911

%	f.t.	E	
0	70.5	-	
3.5	68	-	
5.7	80	68	(1+2)
8.1	90	"	
11.0	100	"	
14.7	110	"	
19.8	120	"	
26.7	130	"	
36.7	140	-	
42.9	145	-	
50.7	150	-	
62.0	155	-	
79.7	160	-	
86.8	161	-	
94.2	160	-	
97.5	163	-	
100	166	-	

Stilbene ($C_{14}H_{12}$) + Antimony trichloride ($SbCl_3$)

Vanstone, 1914

mol%	f.t.	E	
100	73.0	-	
90	65.2	-	
80	46.4	-	
75	83.8	-	(1+2)
70	96.3	-	
67.98	97.3	-	
66.6	97.3	-	
60	88.5	-	
58.58	83.8	75	
50	80.7	75.5	
32.46	100.5	76	
0	125.0	-	

Stilbene ($C_{14}H_{12}$) + Antimony tribromide ($SbBr_3$)

Vanstone, 1914

mol%	f.t.	E	
100	94.9	-	
90.03	89.5	-	
80	81.7	-	
79.20	81.0	-	
70.03	69.0	-	
70	62.0	-	
60	101	-	(1+2)
58.40	102	-	
51.18	94.3	90	
40.69	90.0	90	
34.02	100.0	-	
30.28	104.5	-	
21.36	112.0	-	
10.04	120.0	-	
0	125.0	-	

Diphenylmethane ($C_{13}H_{12}$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	
0	26	-	
7.9	22.5	-	
10.7	30	22.5	
15.1	40	"	
20.2	50	"	
26.0	60	"	
33.0	70	"	
41.6	80	"	
46.8	85	"	
52.7	90	"	
59.8	95	"	
66.3	98	"	
72.9	100	-	(1+2)
78.1	98	-	
82.2	95	67	
86.7	90	"	
89.4	85	"	
91.5	80	"	
93.2	75	"	
95.7	65	-	
97.8	70	67	
100.0	73	-	

Kurnakov, Krotkov and Oksman, 1915 and Kurnakov, 1924

mol%	%	d	η
	100°		
0	0	0.9443	827
50.0	57.38	1.4934	1491
65.13	71.55	.7591	1874
75.0	80.16	.9577	2011
80.0	84.34	2.0715	2013
85.0	88.41	.1907	1984
100	100	.6175	1524

Diphenylmethane ($C_{13}H_{12}$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	
0	26	-	
12.8	22.5	-	
16.9	30	22.5	
22.8	40	"	
29.5	50	"	
37.5	60	"	
47.8	70	"	
53.6	75	"	
60.2	80	"	
67.8	85	"	
73.6	88	"	
81.1	90	-	(1+2)
86.2	88	82	
89.6	85	82	
92.2	82		
93.7	85	82	
96.6	90	82	
100.0	94	-	

s-Diphenylethane ($C_{14}H_{14}$) + Antimony trichloride ($SbCl_3$)

Vanstone, 1914

mol%	f.t.	E	
100	73.0	-	
98.14	71.2	-	
95.03	68.1	66.4	
91.10	69.4	"	(1+4)
86.11	74.6	"	
83.60	76.5	-	
80.96	77.3	-	
78.71	77.2	-	
74.00	75.9	74.2	
73.33	74.2	-	
71.25	74.7	74.2	
67.01	76.0	-	(1+2)
65.90	75.6	-	
57.45	72.9	-	
53.70	71.4	-	
47.12	66.4	43.3	
38.10	59.7	"	
32.10	51.8	"	
23.85	43.3	"	
9.66	48.9	"	
0	51.5	-	

s-Diphenylethane ($C_{14}H_{14}$) + Antimony tribromide ($SbBr_3$)

Vanstone, 1914

mol%	f.t.	E	
100	94.9	-	
89.74	85.1	84.3	
85.00	85.6	84.3	(1+4)
78,84	87.1	-	
70.00	84.3	-	
60.42	79.0	-	
52.58	72.2	42.8	
44.88	63.3	"	
43.09	61.8	"	
36.10	53.0	"	
27.00	43.2	"	
13.96	46.6	"	
3.74	50.5	-	
0	51.5	-	

Triphenylmethane ($C_{19}H_{16}$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	%	f.t.	E
0	92	-	43.5	48.5	-
11.8	85	-	48.1	49.5	-
19.3	80	-	54.5	48.5	-
32.0	70	-	62.8	45	-
42.4	60	-	68.3	40	-
49.6	50	-	72.0	35	-
54.1	42	-	76.6	45	35
			82.4	55	"
			90.6	65	"
			96.1	70	"
(1+1)			100.0	73	-

Kurnakov, Krothov and Oksman, 1915 and Kurnakov, 1924

mol%	%	d	η
	100°		
0	0	1.0195	3224
33.33	31.67	.2629	3450
40.0	38.20	.3320	3512
45.0	43.14	.3884	3458
50.0	48.11	.4449	3388
66.67	64.97	.6938	3241
100	100	2.6175	1524

Triphenylmethane ($C_{19}H_{16}$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	%	f.t.	E
0	92	-	73.3	60	48
18.0	85	-	79.5	70	"
30.1	80	-	86.4	80	"
47.0	70	-	95.2	90	"
58.2	60	48	100.0	94	-
67.1	48	-			

Kordes, 1926

mol%	f.t.
0	92
42	48 E
100	94

Kurnakov, Krotkov and Oksman, 1915 and Kurnakov, 1915

mol%	%	d	η
		75°	
0	0		
25.0	31.95		
40.0	49.56	1.3541	7810
45.0	54.67	1.8150	8892
50.0	59.58	2.0552	9174
60.0	68.86	.2468	9305
66.67	74.67	.3467	9165
70.0	77.47	.5224	8840
75.0	81.56		
100.0	100.00		

%	d	η	d	η
	95°		130°	
0	1.0195	3496	1.0191	3224
31.95	.3346	3819	-	-
49.56	-	-	1.5822	3906
54.67	-	-	.6781	4033
59.58	1.7884	4336	.7832	4159
68.86	.0260	4562	2.0218	4309
74.67	2.2170	4627	.2019	4365
77.47	.3157	4611	-	-
81.56	.4902	4530	2.4805	4303
100.00	3.6926	3309	3.6807	3627

Naphthalene ($C_{10}H_8$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911-12

%	f.t.	E	
0	79.4	-	
15.2	75	59	
26.8	70	"	
35.0	65	"	
42.8	59	-	
48.4	65	59	
53.0	70	"	
58.8	75	"	
65.0	80	"	
74.4	85	"	
78.0	86	-	(1+2)
82.8	85	65	
88.7	80	"	
91.6	75	"	
93.0	70	"	
94.0	65	-	
97.2	70	65	
100	73	-	

Kurnakov, Krotkov and Oksman, 1915 and Kurnakov, 1924

mol%	%	d	η
		80°	
0	0	0.9790	886
5.95	10.0	-	-
12.4	20.0	-	-
25.0	27.08	-	-
32.75	46.26	-	-
50.0	63.87	1.6620	1942
60.0	72.62	-	-
66.67	77.95	1.9488	2311
70.0	80.94	-	-
75.0	84.13	2.1097	2447
77.5	85.89	.1553	2485
80.0	87.61	.2050	2459
85.0	90.92	-	-
90.0	94.07	2.4342	2355
95.0	97.11	-	-
100	100	2.6646	2108

%	d	η	d	η
	90°		150°	
0	0.9696	759	0.9264	217
10.0	1.0349	816	-	-
20.0	.1048	880	-	-
27.08	.2673	1063	-	-
46.26	.3919	1210	-	-
63.87	.6134	1513	1.5675	364
72.62	.8148	1768	-	-
77.95	.9428	1866	1.8971	428
80.94	.9889	1921	-	-
84.13	2.0932	1962	2.0031	448
85.89	-	-	-	-
87.61	2.1938	1989	2.0899	452
90.92	.2957	1933	-	-
94.19	.4057	1937	2.3039	465
97.11	-	-	2.4108	465
100	2.6352	1784	2.4954	459

Naphthalene ($C_{10}H_8$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911-12

%	f.t.	E	
0	79.4	-	
23.7	75	57	
37.5	70	"	
48.6	65	"	
56.8	60	"	
61.2	57	-	
68.0	60	57	(1+2)
74.3	62.5	"	
81.3	65	"	
84.9	66	-	
86.7	65	-	
88.2	70	65	
90.1	75	"	
92.4	80	"	
94.9	85	"	
97.7	90	"	
100.0	94	-	

Naphthalene ($C_{10}H_8$) + Antimony triiodide (SbI_3)

Pushin, 1948

mol%	f.t.	E	mol%	f.t.	E
100	170	-	40	145.5	72
89.5	166	-	30.5	137	73
79.5	163.5	-	20	125	74
67	159	77	10	-	76
49.5	150.5	71	0	80	-

Naphthalene ($C_{10}H_8$) + Titanium tetrachloride ($TiCl_4$)

Pushin, 1948

mol%	f.t.	E	mol%	f.t.	E
100	-23	-	50	55	-
86	+7	-	40	62.5	-26
78.5	20	-26	25	71	-
69.5	34	-	12	76	-
60	46	-	0	80	-

Tetraphenylsilicane ($C_{24}H_{20}Si$) + Tetraphenyllead ($C_{24}H_{20}Pb$)

Pascal, 1912

%	f.t.	E	%	f.t.	E
0	233.0	-	65.00	218.8	218.8
13.04	230.0	226.3	71.43	219.4	218.7
25.92	226.8	223.7	80.00	221.6	220.2
37.50	224.4	220.9	90.91	225.0	222.2
50	221.1	219.3	100.00	227.7	-
62.50	219.1	218.9			

Tetraphenylsilicane ($C_{24}H_{20}Si$) + Tetraphenyltin ($C_{24}H_{20}Sn$)

Pascal, 1912

%	f.t.	E	%	f.t.	E
0	233.0	-	67.27	221.2	221.0
9.09	231.3	229.8	75	222.1	221.0
25.92	228.0	225.2	83.32	223.3	221.1
37.50	225.7	223.1	90	224.6	222.0
50	223.3	222.0	100	225.7	225.7
39.68	222.1	221.2			

Tetraphenylsilicane ($C_{24}H_{20}Si$) + Tetraphenylgermanium ($C_{24}H_{20}Ge$)

Drew and Landquist, 1935

%	m.t.	f.t.
0.0	237.5	237.5
20.0	235.5	237.1
42.5	234.3	236.0
76.7	232.7	234.5
100.0	233.4	233.4

Methylene iodide (CH_2I_2) + Mercuric iodide (HgI_2)

Retgers, 1893

%	f.t.
36.71	180
14.23	100
2.43	15

Chloroform ($CHCl_3$) + Aluminium bromide etherate ($AlBr_3$-$C_4H_{10}O$)

Plotnikov and Kaplan, 1948

mol%	κ	mol%	κ
7.0	4.9	29.0	17.0
8.0	6.1	31.8	17.0
14.1	12.0	33.2	18.0
17.5	14.6	33.2	17.6
17.7	13.1	36.3	17.6
19.1	14.1	39.1	17.7
20.2	14.4	44.0	17.0
25.3	15.6	50.0	16.6
27.6	17.1	58.0	15.0
		60.0	13.6

Chloroform ($CHCl_3$) + Stannic iodide (SnI_4)

Dorfman and Hildebrand, 1927

%	mol%	f.t.
4.94	0.981	10.00
8.28	1.692	25.00
12.91	2.747	40.00

Iodoform (CHI_3) + Antimony tribromide ($SbBr_3$)

Rheinboldt and Schneider, 1928 (fig.)

%	f.t.	E	%	f.t.	E
0	120	-	60	70.5	63
10	115	63	70	71	"
20	108	"	80	80	"
30	100.5	"	90	88.5	"
40	92	"	100	94.5	"
50	82	"			

Carbon tetrachloride (CCl_4) + Titanium tetrachloride ($TiCl_4$)

Sackmann, 1955

mol%	f.t.	m.t.	f.t. II	tr.t. (CCl_4)	E
0	-22.8	-	-	-	-
3	-28	-29.4	-51.5	-	-
5	-27.5	-29.3	-55.3	-	-74
5	-31.8	-35.2	-55.3	-	-
10	-34.7	-43	-	-56.5	-
12.5	-39.5	-45	-	-56	-73.2
15	-43.5	-	-	-56	-73.2
15	-49.3	-	-	-56.5	-70.7
20	-54.5	-	-	-57	-74
25	-60.5	-	-	-	-70
25	-58.5	-	-	-	-72
30	-62	-	-	-	-70
35	-64.5	-	-	-	-72
40	-	-	-	-	-68
45	-65	-	-	-	-70
55	-56	-	-	-	-68.5
65	-49.5	-	-	-	-68.5
75	-44	-	-	-	-70
85	-35.4	-	-	-	-70
95	-29.8	-	-	-	.72.5

Carbon tetrachloride (CCl_4) + Titanium tetrachloride ($TiCl_4$)

Nasu, 1933

mol%	f.t.	E	mol%	f.t.	E
100.0	-24.8	-	52.6	-56.7	-66.6
93.7	-27.6	-66.3	49.7	-58.1	-65.9
91.7	-28.6	-66.3	47.0	-61.1	-66.6
89.9	-31.0	-66.0	38.8	-65.3	-
86.4	-32.4	-66.3	35.6	-62.8	-
81.6	-35.4	-66.0	30.7	-60.3	-66.3
77.4	-38.2	-66.3	26.2	-54.2	-66.3
76.7	-40.0	-66.1	21.1	-50.2	-66.1
68.0	-44.3	-65.8	16.5	-39.4	-66.2
64.0	-47.6	-66.1	15.0	-41.3	-66.3
59.7	-50.7	-65.0	8.1	-33.4	-66.5
56.0	-53.2	-65.9	6.6	-31.4	-66.1
			4.2	-26.9	-67.3
			0.0	-22.9	-

mol% L	mol% V	b.t.
	760mm	
0.00	0.0	76.6
9.76	0.67	79.2
21.16	1.32	84.1
31.27	2.50	89.4
44.27	4.84	93.6
57.00	8.40	100.3
63.44	11.49	106.0
75.81	21.35	116.7
84.89	45.60	124.0
91.64	70.83	129.5
98.03	89.97	134.1
100.00	100.00	135.7

Carbon tetrachloride (CCl_4) + Stannic chloride ($SnCl_4$)

Campbell and Hickman, 1953

mol% L	mol% V	b.t.
16.3	5.5	80
49.5	25.0	90
74.4	49.5	100
93.0	83.2	110

Carbon tetrachloride(CCl_4) + Lead tetrachloride ($PbCl_4$)

Sackmann, 1956

mol%	f.t.	tr.t.	E
0	-22.8	-	-
10	-31	-49.5	-
20	-37	-49.5	-
30	-46	-52.5	-72
35	-50	-	-67
35	-50	-	-
40	-57	-	-62.5
50	-	-	-59
60	-53	-	-57
70	-46.5	-	-59.5
75	-50.5	-	-81 sic
80	-39.5	-	-60.6

Carbon tetrachloride (CCl_4) + Stannic chloride ($SnCl_4$)

Sackmann, 1955

mol%	f.t.	m.t.	tr.t.	E
0	-22.8	-	-	-
2.5	-27.5	-32.5	-53	-
5	-34	-38	-51	-
5	-33	-38.2	-51	-
7.5	-36	-	-54.5	-81
10	-42.5	-	-50.5	-
12.5	-47	-	-52	-77
15	-52.5	-	-	-75.2
20	-55.5	-	-	-75.2
30	-62	-	-	-72
40	-	-	-	-69
45	-68.5	-	-	-70.2
50	-68	-	-	-70.5
60	-56	-	-	-70.5
70	-53.5	-	-	-70.5
80	-47.3	-	-	-70.5
90	-40.5	-	-	-72
95	-35	-	-	-70
100	-32.8			

Carbon tetrachloride (CCl_4) + Tin tetraisopropyl ($SnC_{12}H_{28}$)

Schulze, 1951

mol%	d	mol%	d
	25°		
0	1.58358	48.7882	1.25509
11.3989	.46716	83.0900	.16252
28.1554	.34793	98.7983	.13272
35.8572	.30827	100	.13035

Carbon tetrachloride (CCl_4) + Vanadium tetrachloride (VCl_4)

Whittaker and Yost, 1949

mol %	f.t.	m.t.
0	-23	-
10	-37	-43
20	-46	-64
30	-56	-64
37	-64	-64
40	-63	-64
50	-57	-64
60	-51	-
70	-45	-57
80	-39	-45
90	-32	-34
100	-25.7	-

%	d	χ
	24.5°	
1.000	6.031	1.8255
0.7625	4.689	.7702
0.4983	3.104	.7078
0.2542	1.474	.6490
0.1331	0.5933	.6173
1.000	5.923	.8198
0.7615	4.653	.7695
0.4639	2.849	.6984
0.3188	1.896	.6632
0.1772	0.9043	.6278
0.0876	0.2379	.6056
0.0445	0.0918	.5947

Carbon tetrabromide (CBr_4) + Antimony tribromide. Aluminum bromide ($SbBr_3$).($AlBr_3$)

Gorenbein and Kriss, 1951

%	d	η	ϰ
	85°		
50.5	3.131	4766	24.7
54.4	.145	5420	31.2
60.7	.175	6422	41.2
79.2	.256	11601	66.2
87.6	.283	15315	75.1
89.9	.305	16233	77.8
92.5	.317	18055	-
95.0	.324	19704	81.0
100.0	.346	23466	84.3
	90°		
50.5	3.118	4283	26.8
54.4	.133	4850	34.0
60.7	.162	5614	45.1
79.2	.244	9949	73.0
87.6	.279	12924	83.0
89.9	.292	13637	86.0
92.5	.306	15174	-
95.0	.313	16527	90.0
100.0	.335	19561	93.9
	95°		
50.5	3.106	3899	28.9
54.4	.120	4320	36.8
60.7	.152	5042	48.8
79.2	.231	8735	79.4
87.6	.267	11299	91.1
89.9	.281	11688	94.0
92.5	.294	13048	-
95.0	.301	14077	99.1
100.0	.322	16399	103.4
50.5	3.094	3527	31.6
54.4	.108	3920	40.1
60.7	.135	4539	53.0
79.2	.218	7698	88.4
87.6	.258	9655	102.0
89.9	.270	10192	105.0
92.5	.281	11292	-
95.0	.291	12091	111.3
100.0	.313	13957	116.4

Carbon tetrabromide (CBr_4) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100.0	97.1	50.5	50.2
99.0	96.6	44.1	44.0
97.3	94.5	33.9	44.7
88.5	85.7	32.4	46.1
70.9	72.5	26.3	51.7
70.5	68.5	14.5	68.8
57.5	56.7	9.7	77.6
52.4	52.3	0.0	90.1

Izbekov, 1925

mol%	f.t.	mol%	f.t.
100.0	97.4	50.0	40.5
94.9	92.3	40.1	43.7
89.3	84.9	33.7	48.9
86.2	81.0	26.9	57.1
74.2	66.1	19.1	67.1
69.2	60.2	16.0	71.9
64.1	54.0	8.9	80.1
55.4	45.6	0.0	92.0

Ethyl bromide (C_2H_5Br) + Aluminum bromide ($AlBr_3$)

Vetoup, 1938

mol%	$\varkappa$	mol%	$\varkappa$
4.8	0.5	37.9	33.9
10.0	3.2	40.0	34.3
15.2	9.4	41.0	33.1
23.7	20.5	42.7	33.0
24.0	20.9	43.1	32.9
28.9	26.8	46.2	31.0
30.0	28.5	47.1	30.4
34.2	25.1	51.0	28.5
37.4	32.7	51.2	26.8

Gorenbein, 1940

%	d		
	0°	10°	20°
70.62	2.389	2.357	-
70.55	.375	.358	2.335
60.53	.198	.182	.164
54.86	.116	.095	.076
48.26	.015	1.995	1.978
36.62	1.863	.842	.826
28.11	.767	.747	.727
14.69	.627	.612	.591
5.57	.546	.526	.507
3.99	.527	.506	.489
3.27	.523	.506	.484
2.40	.516	.498	.479
0	.492	.477	.457

%	η		
	0°	10°	20°
70.62	9554.5	6357.4	-
69.92	8383.8	6002.8	4698.2
65.42	5766.6	4467.8	3584.3
62.20	4359.8	3485.9	2875.2
61.76	4175.2	3353.4	2765.6
60.53	3857.3	-	-
58.00	3151.9	2588.7	2183.7
54.86	2756.0	-	1808.0
44.99	1591.4	1363.0	1196.3
40.12	1287.5	1120.7	994.5
36.62	-	-	890.0
35.03	1095.6	972.1	869.5
28.11	-	-	-
22.34	739.2	707.2	646.0
18.53	766.6	695.7	631.4
16.43	-	-	-
11.43	-	-	-
10.73	-	-	-
9.50	-	-	-
8.33	-	-	-
5.57	559.8	501.6	460.3
5.34	-	-	459.4
3.27	531.8	482.0	443.1
2.40	-	476.5	438.7
0	494.6	445.3	410.2

Ethyl bromide (C_2H_5Br) + Antimony tribromide. Aluminum tribromide ($SbBr_3.AlBr_3$)

Gorenbein, 1940

%	d 0°	d 10°	d 20°
71.30	2.601	2.581	2.557
69.5	.551	.529	.509
68.74	.534	.511	.490
65.1	.448	.426	.399
64.99	.447	.425	.405
64.15	.430	.408	.387
61.73	.374	.351	.331
59.66	.330	.308	.287
59.0	.317	.295	.273
57.56	.289	.238	-
54.52	.226	.205	.185
49.9	.122	.102	.082
46.85	.087	.065	.045
41.85	.005	1.983	1.964
38.42	1.949	.929	.909
37.33	.934	.917	.892
33.01	.869	.851	.831
30.41	.839	.818	.798
27.80	.825	.805	.786
22.44	.733	.717	.693
19.90	.705	.685	.675
17.33	.676	.656	.636
12.82	.645	.625	.587
5.31	.550	.531	.511
0	.492	.477	.457

%	η 0°	η 10°	η 20°
71.30	-	6080.6	4483.5
69.50	7086.1	5259.5	4072.4
68.74	-	5142.0	3832.2
65.10	4714.3	3637.7	3637.7
64.99	4596.1	3648.1	2891.3
61.73	3548.6	2998.7	2362.2
59.66	3213.2	2638.8	2153.4
59.00	3048.8	2420.6	2001.2
54.52	1775.6	2032.2	-
49.92	1635.0	1563.7	1376.4
46.85	1635.0	1383.3	1200.7
38.42	1282.2	994.3	870.4
30.41	915.8	800.9	717.0
19.90	709.6	634.5	573.7
17.33	668.1	597.3	541.6
12.82	604.3	552.3	494.4
5.31	541.6	493.1	451.1
0	494.6	445.3	410.2

%	κ 0°	κ 10°	κ 20°
68.74	13.08	15.92	18.97
64.99	13.46	16.15	18.57
61.73	13.77	16.19	18.30
59.66	13.40	15.73	17.42
54.52	-	14.08	-
49.92	9.585	13.17	-
46.85	9.585	10.52	11.41
38.42	6.38	6.64	7.20
30.41	3.64	3.73	3.95
19.90	1.20	1.23	-
17.33	0.898	-	-
15.47	.511	-	-
12.82	.463	-	-
10.03	.303	-	-
9.025	.258	-	-
5.31	.097	-	-

Gorenbein, 1940

%	κ 0°	κ 10°	κ 20°
70.62	20.73	25.17	-
69.92	21.48	26.06	30.47
65.42	24.77	29.05	33.23
62.20	25.75	29.56	33.16
61.76	25.75	29.64	33.25
60.53	25.37	28.88	-
58.00	25.81	28.97	32.12
54.86	25.10	27.51	29.91
48.26	-	-	27.56
44.99	20.10	21.50	23.73
40.12	15.86	16.83	18.12
36.62	12.34	-	13.66
35.03	11.85	12.45	13.46
28.11	6.993	7.257	7.764
22.34	4.142	4.138	4.348
18.53	2.640	2.675	2.878
16.43	1.569	-	-
11.43	0.795	-	-
10.73	-	0.757	0.800
9.50	0.603	-	-
8.33	0.491	-	-
5.57	0.239	-	-
5.34	-	-	-
3.27	0.1364	-	-
2.40	0.0915	0.092	0.091
0	-	-	-

Fairbother and Scott, 1955

25°

%	κ	%	κ
0.3414	0.0084	4.094	0.119
0.9361	.0216	6.246	.214
1.285	.0302	6.899	.245
2.339	.0578	11.44	.585
3.570	.0972	15.90	1.300

Plotnikov, 1902

mol%	κ
7	2.0
16	9.4
23	20.5
39	34.3

Ethyl bromide (C_2H_5Br) + Titanium tetrachloride ($TiCl_4$)

Wertyporoch and Altman, 1934

molarity	κ	molarity	κ
	0°		
0.092	0.282	0.647	0.285
.226	.283	.842	.357
.441	.303		

Ethyl bromide (C_2H_5Br) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altman, 1934

molarity	κ	molarity	κ
	0°		
0.0436	0.01096	0.4160	0.01640
.0868	.01360	0.7950	.01345
.1293	.01550	1.1400	.01120
.2550	.01555	1.4560	.0973

Ethyl bromide (C_2H_5Br) + Tantalum pentachloride ($TaCl_5$)

Namoradze and Zvyagintsev, 1939

g/100cc	f.t.
0	11.14
20	17.43

Ethyl iodide (C_2H_5I) + Aluminum bromide ($AlBr_3$)

Katznelson, 1938

mol%	κ
19.0	54.2
30.4	72.3
40.1	57.6

Ethylene bromide ($C_2H_4Br_2$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1910

%	f.t.	%	f.t.
0	10	56.0	40
11.5	6	63.7	50
21.3	+2	71.5	60
29.7	-2	79.1	70
36.1	+10	86.8	80
42.1	20	94.5	90
48.7	30	100	96

Ethylene bromide ($C_2H_4Br_2$) + Antimony tribromide ($SbBr_3$)

Gorenbein, 1954

%	D f.t.
10.63	-3.69
12.35	4.30
14.66	5.09
16.04	5.53

Ethylene bromide ($C_2H_4Br_2$) + Stannic iodide (SnI_4)

Dorfman and Hildebrand, 1927

%	mol%	f.t.
9.11	2.915	10.00
14.16	4.714	25.00
20.91	7.347	40.00
100	100	143.50

Tetrachlorethane ($C_2H_2Cl_4$) + Titanium tetrachloride ($TiCl_4$)

De Pauw, 1922

%	d	%	d
	25°		
0	1.58909	15.18	1.69672
8.36	.64638	19.66	.73168
12.20	.69127	28.81	.80897
14.56	.69193		

Butta, 1956.

mol % L	mol % V	b.t.
0	0	146.1
21.4	29.2	141.4
29.5	38.0	140.2
42.9	50.3	138.6
49.8	56.0	137.8
52.1	57.9	137.5
66.5	70.5	136.4
79.4	80.0	135.7
91.3	91.1	135.4
100.0	100.0	136.6

Propyl chloride (C_3H_7Cl) + Titanium tetrachloride ($TiCl_4$)

Wertyporoch and Altmann, 1934

molarity	κ	molarity	κ
	0°		
0.1265	1.1116	0.8420	2.032
.2260	2.270	1.4940	1.765
.4410	2.622	2.3080	1.167

Propyl chloride (C_3H_7Cl) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altmann, 1934

molarity	κ	molarity	κ
	0°		
0.2550	0.578	0.7950	0.523
.3360	.577	1.1400	.516
.4160	.575	1.4560	.435
.6110	.577		

Isopropyl chloride (C_3H_7Cl) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altmann, 1934

molarity	κ	molarity	κ
	0°		
0.0514	4.160	0.4975	4.730
.1160	4.850	.9270	4.625
.2502	5.310	1.6840	3.250

Allyl chloride (C_3H_5Cl) + Titanium tetrachloride ($TiCl_4$)

Wertyporoch and Altmann, 1934

molarity	κ	molarity	κ
	0°		
0.0461	0.0754	0.2700	0.00381
0.0920	.00908	0.4410	0.00405
0.1820	.00377	0.8420	0.00556

Allyl chloride (C_3H_5Cl) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altmann, 1934

molarity	κ	molarity	κ
	0°		
0.0436	0.01800	0.4160	0.01430
.0868	.01900	.6110	.01235
.1715	.01625	.7950	.01140

Perfluorpentane (C_5F_{12}) + Tungsten hexafluoride (WF_6)

Barber and Cady, 1951

mol% V	mol% L	p	mol% V	mol% L	p
		25°			
100	100.0	1019.4	48.7	34.2	852.1
94.6	94.8	1021.3	47.8	33.2	842.6
90.3	89.6	1020.9	34.6	22.2	790.4
83.7	81.8	1012.9	32.5	19.7	778.5
84.1	81.6	1013.9	16.8	9.3	706.8
68.1	61.4	950.5	18.8	-	709.6
-	47.6	909.4	0.0	0.0	646.5
60.3	45.8	902.2			

mol% V	mol% L	b.t.	mol% V	mol% L	b.t.
		1140mm			
100.0	100.0	28.15	-	35.5	33.48
94.6	94.7	28.11	47.2	32.0	33.77
94.7	94.8	28.11	-	22.6	35.54
89.8	89.1	28.14	-	22.0	35.71
83.8	81.8	28.33	34.2	21.8	36.01
66.2	57.8	30.22	15.3	9.2	38.66
59.4	47.7	31.52	0.0	0.0	40.86
-	46.0	31.74			

Perfluorheptane (C_7F_{16}) + Stannic chloride ($SnCl_4$)

Campbell and Hickman, 1953 (fig.)

mol%	sat.t	mol%	sat.t.
95	65	40	86
80	95	20	60
64	97	15	35
60	96		

Perfluorcyclopentane (C_5F_{10}) + Tungsten hexafluoride (WF_6)

Rohrbach and Cady, 1951

mol % L	mol % V	p
	25°	
100.0	100.0	1019.4
93.9	93.2	1030.3
89.6	88.9	1035.0
82.4	82.6	1030.7
52.5	57.6	1004.4
40.0	45.5	978.1
25.9	34.3	944.3
17.9	25.5	919.2
11.7	18.0	896.9
0.0	0.0	833.4
	45°	
100.0	100.0	1982.4
94.4	93.9	1998.0
93.7	-	2000.5
89.7	89.5	2008.8
-	83.1	2009.0
82.6	82.7	2010.3
75.2	76.6	2004.4
57.3	60.6	1967.1
40.5	44.6	1899.1
26.2	33.0	1834.9
17.9	23.9	1788.3
0.0	0.0	1642.3

Fluorbenzene (C_6H_5F) + Aluminum bromide ($AlBr_3$)

Wertyporoch and Adamus, 1934

molarity		ϰ
	0°	
0.224		0.0377
.360		.0965
.546		.2165
.708		.5090

Fluorbenzene (C_6H_5F) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	tr.t.
0	-39.2	-	-
2.4	-40.5	-	-
5.2	-35	-40.5	-
11.0	-25	-40.5	-
17.3	-15	"	-
21.4	-10	"	-
26.4	-5	"	-
34.1	0	"	-
45.8	5.5	-	-
54.0	8	-	-
70.2	10	-	-
			(1+1)
53.6	15	-	5.5
61.6	25	-	5.5
69.6	35	-	-
77.7	45	-	-
85.7	55	-	-
93.8	65	-	-
100	73	-	-

Fluorbenzene (C_6H_5F) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	%	f.t.
0	-39.2	-	35.3	45
1.3	-39.5	-	45.5	55
4.3	-25	-39.5	60.8	65
6.7	-15	"	72.8	70
9.4	-5	-	81.8	75
12.6	+5	-	88.3	80
16.8	15	-	93.5	85
21.8	25	-	97.6	90
27.8	35	-	100	94

Chlorbenzene (C_6H_5Cl) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	
0	-45.2	-	
4.3	-47	-	
7.0	-40	-47	(1+1)
11.1	-30	"	
16.5	-20	"	
20.5	-15	"	
25.3	-10	-	
32.5	-5	-	
44.2	0	-	
58.5	+4	-	
49.3	10	-	(0+1)
56.0	20	-	
64.0	30	-	
72.1	40	-	
80.1	50	-	
88.2	60	-	
97.1	70	-	
100.0	73	-	

Chlorbenzene (C_5H_5Cl) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.
0	-45.2	-	36.8	30
5.2	-47	-	45.4	40
6.8	-40	-47	55.6	50
9.6	-30	"	65.8	60
12.6	-20	"	76.0	70
16.0	-10	"	86.3	80
20.0	1	"	96.7	90
24.6	+10	-	100.0	94
30.0	20	-		

Chlorbenzene (C_6H_5Cl) + Stannic chloride ($SnCl_4$)

Bingham, 1907

C.S.T. = -15°

Brombenzene (C_6H_5Br) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	tr.t.	
0	-31	-	-	
4.8	-32.5	-	-	
9.0	-35	-	-	
6.8	-30	-32.5	-	
10.7	-25	"	-	
14.8	-20	"	-	
19.2	-15	"	-	
23.9	-10	"	-	(1+1)
28.7	-5	"	-	
34.3	0	"	-	
40.3	+5	"	-	
46.0	3	"	-	
51.0	5	"	-	
54.5	6	-	-	
59.1	7.0	-	-	
45.0	10	-	3	(0+1)
52.0	20	-	3	
59.8	30	-	3	
68.0	40	-	3	
76.6	50	-	3	
85.8	60	-	3	
90.7	65	-	3	
96.4	70	-	3	
100.0	73	-	-	

Brombenzene (C_6H_5Br) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.
0	-31	-	46.8	35
5.7	-32	-	54.6	45
9.5	-25	-32	63.0	55
15.0	-15	"	71.9	65
20.8	-5	"	81.2	75
26.8	+5	"	90.7	85
33.0	15	-	95.6	90
39.6	25	-	100.0	94

Iodobenzene (C_6H_5I) + Aluminum bromide ($AlBr_3$)

Wertyporoch and Adamus, 1934

m	κ
0°	
0.141	0.044
0.291	0.1085
0.387	0.1885
0.550	0.220
0.798	0.904
0.902	1.015

Iodobenzene (C_6H_5I) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	tr.t.	
0	-28.6	-	-	
3.0	-30	-45	-	
12.8	-35	"	-	
22.2	-40	"	-	
29.8	-45	-	-	
11.7	-34.5	-	-	(1+1)
17.8	-25	-	-34.5	
26.4	-15	-	"	
40.8	-5	-	-34	
49.1	-3	-	"	
32.5	-35	-45	-	(0+1)
35.6	-25	"	-	
38.9	-15	"	-	
42.4	-5	"	-	
46.4	+5	"	-	
50.7	15	"	-	
56.0	25	"	-	
62.3	35	"	-	
69.6	45	"	-	
78.0	55	"	-	
88.8	65	-	-	
95.7	70	-	-	
100.0	73	-	-	

Iodobenzene (C_6H_5I) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.	E
0	-28.6	-	51.3	30	-
7.0	-30.3	-32	57.6	40	-
14.3	-32	-	64.2	50	-
21.6	-20	-32	71.1	60	-
27.5	-10	"	78.4	70	-
33.4	0	"	86.3	80	-
39.3	+10	"	95.3	90	-
45.2	20	-	100.0	94	-

o-Dichlorbenzene ($C_6H_4Cl_2$) + Aluminum bromide ($AlBr_3$)

Wertyporoch and Adamus, 1934

molarity	κ	molarity	κ
	0°		
0.022	0.063	0.148	0.586
.037	.087	.235	.593
.054	.088	.362	.666
.080	.315	.505	.687
.125	.482	710	.731

p-Dichlorbenzene ($C_6H_4Cl_2$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.	E
0	54.5	-	67.8	50	39.5
14.0	50	39.5	75.7	55	"
30.0	45	"	83.0	60	"
48.0	40	"	89.8	65	"
50.5	39.5	-	96.2	70	"
59.5	45	39.5	100.0	73	-

p-Dichlorbenzene ($C_6H_4Cl_2$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.	E
0	54.5	-	58.8	70	48.5
14.0	51.5	48.5	67.2	75	"
26.5	48.5	-	75.8	80	"
35.9	55	48.5	84.5	85	"
43.1	60	"	93.4	90	"
50.7	65	"	100.0	94	-

p-Dibrombenzene ($C_6H_4Br_2$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.	E
0	88	-	59.0	55	49.5
5.7	85	49.5	64.0	49.5	-
15.4	80	"	71.8	55	49.5
25.2	75	"	79.3	60	"
35	70	"	86.8	65	"
44.5	65	"	95.0	70	"
52.8	60	"	100.0	73	-

p-Dibrombenzene ($C_6H_4Br_2$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.	E
0	88	-	75.3	75	65
10.0	85	65	81.8	80	"
25.2	80	"	88.3	85	"
39.2	75	"	94.3	90	-
52.0	70	"	97.0	92	-
62.2	65	-	100.0	94	-
68.7	70	65			

o-Chlortoluene (C_7H_7Cl) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	%	f.t.	E
0	-36.2	-	47.9	-0.5	-
6.9	-37.5	-	53.1	10	-
10.8	-30	-37.5	58.2	20	-
18.3	-20	"	64.6	30	-
29.2	-10	"	71.8	40	-
37.1	-5	" (1+1)	79.7	50	-
47.9	-0.5	-	88.4	60	-
58.5	+2	-	97.2	70	-
64.2	3	-	100.0	73	-

o-Chlortoluene (C_7H_7Cl) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	%	f.t.	E
0	-36.2	-	38.8	30	-
10.7	-38.5	-	46.8	40	-
12.6	-30	-38.5	56.0	50	-
15.4	-20	"	66.5	60	-
18.7	-10	"	77.8	70	-
22.5	0	-	88.2	80	-
27.0	+10	-	97.0	90	-
			100.0	94	-

m-Chlortoluene (C_7H_7Cl) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	%	f.t.	E
0	-47.8	-	40.0	-14	(0+1)
6.9	-49	-	46.1	0	-
12.3	-40	-49(1+1)	51.6	10	-
20.1	-30	"	57.4	20	-
31.0	-20	"	64.8	30	-
40.0	-49	"	72.8	40	-
51.9	-10	"	80.9	50	-
56.0	-9	-	89.1	60	-
			97.4	70	-
			100.0	73	-

m-Chlortoluene (C_7H_7Cl) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	%	f.t.
0	-47.8	-	37.5	30
8.1	-50	-	45.1	40
9.7	-40	-50	54.4	50
11.7	-30	"	65.0	60
14.3	-20	"	77.0	70
17.5	-10	"	88.2	80
21.3	0	-	97.0	90
25.8	+10	-	100.0	94
31.2	20	-		

p-Chlortoluene (C_7H_7Cl) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	%	f.t.
0	6.2	58.0	20
12.7	3	64.8	30
23.1	0	72.3	40
32.2	-3	80.2	50
39.7	-6	88.8	60
43.8	-7.5	97.4	70
47.2	0	100.0	73
52.2	+10		

p-Chlortoluene (C_7H_7Cl) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	%	f.t.	E
0	6.2	-	56.3	50	2.5
12.3	4.2	-	66.7	60	"
23.3	2.5	-	77.8	70	-
27.5	10	2.5	88.2	80	-
33.0	30	"	97.0	90	-
39.3	30	"	100.0	94	-
47.2	40	"			

Benzal chloride($C_7H_6Cl_2$) + Titanium tetrachloride ($TiCl_4$)

Wertyporoch and Altmann, 1934

molarity	κ	molarity	κ
	0°		
0.0461	0.00658	0.8420	0.05700
.0920	.01314	1.2120	.06770
.2700	.03150	1.4940	.07510
.4410	.04430		

1-Chlornaphthalene($C_{10}H_7Cl$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911-12

%	f.t.	E	
0	-17	-	
8.1	-21	-	
11.1	-10	-21	(1+1)
14.4	0	"	
18.7	+10	"	
24.6	20	"	
28.5	25	-	
33.5	30	-	
39.8	35	-	
47.7	40	-	
53.5	42.5	-	
61.5	45	-	
67.3	46	-	
75.0	46	-	
78.1	50	46	
82.2	55	"	
91.3	60	"	
96.5	70	"	
100	73	-	

1- Chlornaphthalene ($C_{10}H_7Cl$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911-12

%	f.t.	%	f.t.
0	-17	46.7	30
13.8	-21	53.8	40
22.6	-24.5	61.6	50
23.9	-20	69.9	60
27.3	-10	78.6	70
31.2	0	87.5	80
35.5	+10	96.6	90
40.5	20	100.0	94

2-Chlornaphthalene ($C_{10}H_7Cl$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911-12

%	f.t.	tr.t.
0	56	-
16.6	50	25
27.2	45	"
35.4	40	"
41.8	35	"
47.3	30	"
52.3	25	-
56.6	20	-
59.1	16.5	-
52.3	25	- (1+1)
55.6	28.5	25 "
58.2	29.5	- "
61.1	29	28 "
64.0	28	- "
60.2	20	-
62.5	25	-
65.2	30	28
68.3	35	"
71.7	40	"
75.3	45	"
79.1	50	"
83.2	55	"
87.5	60	"
92.2	65	"
97.2	70	"
100.0	73	-

2-Chlornaphthalene ($C_{10}H_7Cl$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911-12

%	f.t.	E
0	56	-
15.5	53	37.5
26.1	50	"
38.5	45	"
49.0	40	"
53.6	37.5	-
55.3	40	37.5
58.8	45	"
62.7	50	"
66.8	55	"
71.0	60	"
75.2	65	"
79.5	70	"
83.8	75	"
88.1	80	"
92.4	85	"
96.7	90	"
100	94	-

1-Bromnaphthalene ($C_{10}H_7Br$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911-12

%	f.t.	E
0	3	-
8.3	- 1	-
12.8	+10	-1 (1+1)
15.7	20	-1
19.0	20	-1
24.0	25	-1
30.3	30	-1
38.5	33	-1
52.4	34.5	-
62.1	33	31.5
64.7	31.5	-
69.7	40	31.5
76.2	50	"
80.2	55	"
84.5	60	"
89.1	65	"
94.8	70	"
100.0	73	-

1-Bromnaphthalene ($C_{10}H_7Br$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911-12

%	f.t.	E
0	+ 3	-
15.8	0	-3.5
31.4	- 3.5	-
34.1	+ 5	-
38.7	15	-
43.9	25	-
49.9	35	-
56.9	45	-
64.7	55	-
72.9	65	-
81.8	75	-
86.3	80	-
90.8	85	-
95.4	90	-
100.0	94	-

Carbon dioxide (CO_2) + Lithium oxide (Li_2O)

Lebeau, 1905

t	p dissoc.	t	p dissoc.
580	0	830	34
610	1	840	37
620	3	860	41
640	8	890	51
710	16	930	61
740	19	1000	91
770	23	1200	300
800	29		

Carbon dioxide (CO_2) + Sodium oxide (Na_2O)

Lebeau, 1903-05

t	p dissoc.	t	p dissoc.
700	1	1050	16
730	1.5	1080	19
820	2.5	1100	21
980	10	1150	28
990	12	1180	38
1010	14	1200	41

Carbon dioxide (CO_2,H_2O) + Sodium carbonate (Na_2CO_3)

Lescoeur, 1892

t	p dissoc.	t	p dissoc.
55	19	80	70
60	25	90	125
70	43	100	310

Carbon dioxide (CO_2) + Potassium oxide (K_2O)

Lebeau, 1903-05

t	p dissoc.	t	p dissoc.
730	0	960	5
790	0.5	970	9
810	1	1000	12
890	3	1090	17

Carbon dioxide (CO_2,H_2O) + Potassium carbonate (K_2CO_3)

Lescoeur, 1892

t	p dissoc.	t	p dissoc.
85	25	110	100
90	36	120	150
100	65	127	198

Caven and Sand, 1914

t	p dissoc.	t	p dissoc.
63.7	4.7	127.4	195
76.3	11.0	137.7	320
90.2	26.1	138.4	330
92.5	29.9	146.3	473
103.5	56.0	147.8	506
104.6	59.6	151.8	604
116.4	112	153.4	647
119.1	129	155.4	706
127.2	193	156.0	724

Carbon dioxide (CO_2)$_2$ + Potassium oxide (K_2O)

Lescoeur, 1896

t	p dissoc.
100	5
120	17
140	69
150	111

Carbon dioxide (CO_2) + Rubidium oxide (Rb_2O)

Lebeau, 1903-05

t	p dissoc.	t	p dissoc.
690	0	900	10
740	2	990	18
830	6	1020	20
870	8	1080	33

Carbon dioxide (CO_2,H_2O) + Rubidium carbonate (Rb_2CO_3)

Caven and Sand, 1914

t	p dissoc.	t	p dissoc.
12.7	0	136.5	163
15.0	0	137.3	171
61.5	0.7	143.2	240
91.2	8.0	146.8	295
95.8	11.3	147.1	300
97.3	12.6	151.5	383
106.5	24.1	151.9	392
109.3	29.2	153.5	427
111.4	33.6	158.4	556
112.9	37.2	158.6	562
120.1	59.4	160	605
120.5	60.9	161	638
121.2	63.8	164	747
135.2	151	170.6	1045
135.3	152		

Carbon dioxide (CO_2) + Cesium oxide (Cs_2O)

Lebeau, 1903 - 1905

t	p dissoc.	t	p dissoc.
610	2	1000	44
680	4	1050	63
805	6	1090	90
860	8	1130	121
890	12	1150	137
980	32	1180	157

Carbon dioxide (CO_2,H_2O) + Cesium carbonate (Cs_2CO_3)

Caven and Sand, 1914

t	p dissoc.	t	p dissoc.
89.5	0.4	153.1	140
103.0	1.5	157.8	202
103.8	1.6	158.1	207
116.5	5.7	160.1	242
117.9	6.5	163.0	303
133.1	24.0	166.1	382
133.6	27.3	169.9	507
135.3	31.6	172.2	601
140.0	47.4	172.7	624
142.7	59.6	175.4	759
144.5	69.2	177.0	851
151.6	124	178.1	921
152.4	132	179.8	1038

Carbon dioxide (CO_2) + Magnesium oxide (MgO)

Marc and Simek, 1913

t	p dissoc.	t	p dissoc.
398.3	0.825	452.4	3.874
412.7	1.380	456.7	.517
424.7	.650	470.0	4.391
426.7	.856	473.7	.578
429.1	.730	481.1	6.162
431.9	.816	489.3	.904
437.9	3.277	490.7	.943
440.6	.362	497.4	10.560
440.7	.059	497.4	9.201
442.5	.153	502.5	9.200
443.5	2.900	508.9	11.169
451.3	3.118	509.1	11.064

Carbon dioxide (CO_2) + Calcium oxide (CaO)

Le Chatelier, 1886

t	p dissoc.	t	p dissoc.
547	27	745	289
610	46	810	678
625	56	812	763
740	255	865	1333

Zavriev, 1909

t	p dissoc.	t	p dissoc.
360	1	825	270
445	8.5	840	342
525	18	860	420
725	71	870	500
750	100	890	610
760	110	892	626
793	170	910	755
795	170	912	791.5
815	230	926	1022

Johnston, 1910

t	p dissoc.	t	p dissoc.
587	1.0	749	72
605	2.3	777	105
631	4.0	786	134
		788	138
671	13.5	795	150
673	14.5	800	183
680	15.8	819	235
682	16.7	830	255
691	19.0	840	311
701	23.0	842	335
703	25.5	852	381
711	32.7	857	420
		871	537
727	44	876	557
736	54	881	603
743	60	891	684
748	70	894	716

Riesenfeld, 1910

t	p dissoc.	t	p dissoc.
700	59	850	370
750	99	900	700
800	195		

Andrusov, 1925

t	p dissoc.	t	p dissoc.
506	0.2	725.5	52.0
553	0.7	760	100.4
596	2.9	800	208
632	6.3	850	463
672	17.3	882	760
704	33.1	902	1008

Smyth and Adams, 1923

t	p dissoc.	t	p dissoc.
C_1+C_2+V		C_1+L+V	
842.3	343.0	1244.5	31528
852.9	398.6	1244.9	31528
854.5	404.1	1267.2	39690
868.9	510.9	1270.1	39535
904.3	879.0	1275.6	53505
906.5	875.0	1278.1	49875
937.0	1350	1296.0	107262
937.0	1340	1296.9	106228
990.6	2053	1304.8	108296
1038.2	3604	1306.4	157411
1049.3	4894	1314.0	193800
1082.5	6758	1324.2	449160
1157.7	14202	1325.0	426360
1226.3	26093	1334.7	798760
1226.2	26093	1336.1	779000
1240.9	29718 E	1338.9	778000
1238.9	30579 E		

Hill and Winter, 1956

t	p dissoc. (1+1)	t	p dissoc. (1+1)
449	0.0170	698	23.10
465	0.0330	706	29.04
466	0.0334	719	37.85
480	0.0483	737	53.80
493	0.0806	747	65.19
508	0.128	767	90.94
528	0.262	767	95.94
540	0.360	784	132.7
540	0.432	785	117.5 ?
554	0.579	803	168.4
556	0.598	816	219.0
564	0.801	822	220.0
579	1.120	832	271.9
596	2.025	850	365.6
607	2.46	853	372.0
615	3.15	864	457.0
632	5.26	870	495.0
640	6.75	876	518.8
656	9.47	886	584.0
662	10.22	898	692.5
673	13.46	904	813.0
686	16.95		

Carbone dioxide (CO_2) + Strontium oxide (SrO)

Jones and Becker, 1927

t	p diss.(1+1)	t	p diss.(1+1)
689	0.1	975	21
708	0.3	984	25
756	0.5	1020	42
775	0.6	1027	47.5
835	1.7	1059	72.5
846	2.2	1070	84
869	3.6	1081	95
913	8.8	1090	111
958	16.0	1103	123

Phosgene ($COCl_2$) + Aluminum chloride ($AlCl_3$)

Germann and Mc Intyre, 1925

%	p	%	p	%	p
250°		25°		0°	
0	1406	24.6	1200	0	555
1.22	1397	28.0	1139	1.20	551
2.49	1380	33.1	1052	2.41	547
4.86	1364	36.1	991	4.56	541
10.90	1308	38.65	931	9.52	524
17.86	1263	41.8	842	21.64	473
27.97	1105	43.7	786	29.18	429
31.82	1087	46.3	703	33.36	398
37.15	888	47.7	658	37.73	343
38.91	945	49.4	578	41.94	281
41.40	756	50.8	522		
44.50	600	52.4	445		
45.31	776	53.7	380		
47.7	457				
50.05	578				
55.81	323				

Germann, 1925

%	d 25°	d 0°	%	d 25°	d 0°
0	1.3685	1.4275	20.89	-	1.5303
2.97	-	1.4422	21.55	1.4825	-
2.99	1.3861	-	30.38	-	1.5760
5.72	-	1.4573	30.50	1.5283	-
5.96	1.4009	-	46.33	-	1.6473
9.70	-	1.4758	46.53	1.6089	-
9.77	1.4215	-			

%	κ 25°	κ 0°	κ -45°
0.00000	0.000218	0.000194	0.000236
.01725	.000333	.000325	.000368
.0357	.000564	.000605	.000676
.0763	.000706	.000852	.001172
.123	.001047	.00130	.00193
.245	.00276	.002935	.00373
1.240	.00604	.00735	.01008
.25	.00416	.00491	.00625
.355	.00717	-	-
2.18	.0180	0.0209	0.0227
2.66	.0181	.0203	.0220
4.62	.098	.1023	.0976
5.00	.0838	.0895	.0806
6.19	.1457	.1444	.1126
10.00	.436	.408	.304
17.3	1.220	.971	.619
17.8	1.53	1.182	.657
29.5	3.35	2.28	-
44.6	5.20	2.77	-
53.0	5.54	-	-

Phosgene ($COCl_2$) + Sodium chloraluminate ($NaAlCl_4$)

Germann and Birosel, 1925

%	p dissoc. 25°	%	p dissoc. 25°	
0	1405	35.3	1305	
8.95	1400	37.9	1270	
9.44	1399.5	40.7	1287	
10.33	1398.5	44.2	1286	
11.08	1398	48.9	1285	
13.06	1397	58.9	1282	no complex
14.47	1396.5	67.5	1276	
17.4	1393.5	79.0	1268	
20.7	1389	95.1	1221	
24.7	1380.5	98.7	742	
28.7	1362	98.9	429	
30.6	1349	99.6	102	
32.8	1331	100.0	9	

Phosgene ($COCl_2$) + Barium chloraluminate ($BaAl_2Cl_8$)

Germann and Birosel, 1925

%	p dissoc.
1st series	
0	1405
9.10	1405 (L_1+L_2)
9.6	1405
10.3	1405
20.3	1405
45.1	1402 (L)
47.2	1401
50.4	1372
52.9	1321
53.1	1192
54.8	1343 (8+3)
57.3	1311
59.8	1241
63.1	1191
64.3	1024
64.5	741
68.3	557 (1+1)
76.4	611
82.5	550
83.4	204
92.2	282
94.9	187
98.0	135
99.7	42
100.0	0
2nd series	
0	1405
36.3	1401 (8+3)
41.7	1401
49.0	1401
54.0	1316
57.0	1232
58.4	1338 (1+1)
60.4	1305
64.1	1130
64.5	731
67.6	611
78.9	624
82.0	498
83.2	327
87.4	293
93.1	258
96.6	148
100.0	0

Next Page

Phosgene ($COCl_2$) + Calcium chloraluminate ($CaAl_2Cl_8$)

Germann and Timpany, 1925

%	p	%	p	
25°				
3.06	1400	59.40	1153	
11.24	1385	62.40	1017	
22.04	1358	63.60	908	
28.01	1320	64.10	767	
		64.80	602	
11.43	1380	65.40	439	
12.70	1376	65.75	272	
15.48	1374	66.45	103	(2+1)
18.90	1365	64.60	73	
22.30	1354	67.25	51	anh
26.00	1339	66.40	43	
30.18	1304	69.8	39	
36.00	1255	80.2	36	
34.35	1228	90.3	8	
37.90	1250	94.8	4	
37.48	1155	100	0	
40.20	1247			
43.00	1246			
45.80	1242			
52.10	1215			
0°				
2.99	554	28.56	517	
10.28	547	29.77	514	
10.60	549	31.20	507	
11.50	546	32.15	495	
13.73	544	32.66	498	
16.39	542	34.24	498	
19.17	540	36.71	483	
21.54	537	38.10	490	
22.95	530	42.95	493	
24.62	531	47.40	484	

%	κ	%	κ
25°		0°	
0.017	0.00106	0.017	0.00100
.034	.001057	.033	.000967
.064	.001280	.062	.001208
.142	.001853	.141	.001428
.276	.001880	.276	.001710
.568	.002560	.564	.002258
1.183	.006217	1.81	.005270
2.37	.02350	2.36	.02230
5.06	.1986	4.96	.1813
6.70	.5815	6.70	.6378
8.73	1.251	8.70	1.203
13.08	5.695	14.68	6.318
14.78	7.550	14.81	7.780
20.37	20.47	17.60	10.250
25.00	31.42	19.51	13.180
27.70	37.85	20.23	15.680
30.64	43.55	24.77	22.800
33.49	47.88	27.43	26.110
36.80	51.05	30.30	29.120
		33.07	30.670
		36.30	30.000

%	d 25°	d 0°
0	1.3696	1.4267
6.03	-	1.4638
6.23	1.4091	-
12.10	-	1.5010
12.19	1.4470	-
12.44	-	1.5037
12.55	1.4528	-
20.70	-	1.5620
21.32	1.5150	-
30.52	-	1.6199
30.73	1.5748	-
20.10	-	1.5550
20.26	1.5055	-

Phosgene ($COCl_2$) +Strontium chloraluminate ($SrAl_2Cl_8$)

Germann and Birosel, 1925

%	p dissoc.	
25°		
0	1405	
16.2	1398.5	
17.4	1395.5	
18.9	1395.5	
22.0	1395.5	
23.4	1393	
24.9	1391	
27.5	1385.5	
30.6	1375	
35.2	1348.5	
30.3	1297	
42.2	1248	
46.2	1151	
48.7	1079	
50.9	1008	(9+5)
53.3	925	
55.5	849	
57.6	756	
59.3	696	
61.0	632	
63.2	942	
64.7	934	
67.1	913	
69.4	853	
70.4	608	
70.8	318	
75.0	539	
78.6	541	
80.0	321	
80.3	147	
80.5	47	
80.6	14	
81.0	0	
85.5	176	(1+1)
87.2	166	
89.2	174	
91.6	170	
94.5	153	
97.0	147	
99.5	116	
100.0	0	

Previous Page

Carbon disulfide (CS_2) + Stannic iodide (SnI_4)

Arctowski, 1896

%	f.t.
16.27	-58
10.22	-84
9.68	-89
9.65	-94
9.4	-114.5

Dorfman and Hildebrand, 1927

%	mol%	f.t.
49.01	10.46	10.00
58.53	14.64	25.00
67.56	20.20	40.00

Doane and Drickamer, 1955

mol %	P at f.t.
25°	
14.61	1
7.99	1000
5.64	2000
3.45	3600
1.96	5000
0.77	7000
0.11	10000

Carbon disulfide (CS_2) + Aluminum bromide ($AlBr_3$)

Kaveler and Monroe, 1928

mol%	f.t.	mol%	f.t.
17.0	0.10	64.4	60.0
21.6	8.4	73.1	70.0
25.8	15.0	74.2	71.0
28.9	20.0	79.1	76.0
32.3	25.0	83.3	81.0
36.2	30.0	87.8	85.0
54.5	50.1		

Plotnikov, Sheka and Yankelevich, 1939

%	d	ε	%	d	ε
0.00	1.2639	2.641	27.24	1.4891	2.792
4.01	.2920	.660	34.88	.5658	.816
10.99	.3454	.706	40.47	.6300	.854
18.96	.4088	.747	45.46	.6884	.881
22.07	.4376	.756	49.07	.7336	.910

Carbon disulfide (CS_2) + Aluminum bromide etherate ($AlBr_3.C_4H_{10}O$)

Plotnikov and Kaplan, 1948

mol%	κ	mol%	κ
5.5	2.3	35.5	27.0
10.5	8.1	39.0	28.2
12.8	12.4	42.1	28.7
16.2	16.4	42.7	28.5
21.0	22.5	43.4	28.0
23.0	23.4	57.3	27.0

Carbon disulfide (CS_2) + Titanium tetrachloride ($TiCl_4$)

Bond and Stephens, 1929

%	f.t.	m.t.	%	f.t.	m.t.
100.00	-	-24.4	43.69	11.0	-
98.96	-	-26.0	37.67	9.6	-
97.47	-	-27.7	33.87	8.3°	-
96.37	-	-29.1	30.47	6.5	-
96.31	-42.2	-28.9	26.17	3.5	-
94.95	-33.7	-30.1	23.56	1.5	-
90.65	-13.7	-	21.09	-1.5	-
88.66	-6.3	-	19.92	-3.3	-
85.35	-0.3	-	15.02	-11.2	-
85.16	+0.1	-	14.44	-12.2	-
82.43	4.0	-	12.26	-16.8	-
79.35	6.8	-	9.53	-23.7	-
73.26	10.2	-	8.23	-27.9	-
68.08	11.5	-	7.05	-31.4	-31.4
65.40	11.8	-	6.44	-36.1	-32.7
61.80	11.9	-	4.15	-	-36.1
60.46	12.0	-	3.41	-	-39.3
58.40	12.0	-	2.86	-	-46.5
58.05	12.0	-	2.01	-	-54.2
56.23	11.9	-	1.64	-	-57.8
52.82	11.8	-	0.96	-	-66.8
48.20	11.5	-			

Carbon disulfide (CS_2) + Stannic chloride ($SnCl_4$)

Schwers, 1912

%	t	d
0.00	21.0	1.26037
"	34.2	.24069
13.8282	25.9	.33067
"	34.5	.31700
27.2410	25.2	.41697
"	34.2	.40184
38.605	24.5	.50003
"	34.3	.48290
44.3085	25.9	.54270
"	34.45	.52696
66.3864	26.5	.74819
"	34.6	.73114
80.7152	25.1	.91772
"	34.5	.89828
100	23.3	2.21355
	58.0	.12436

%	t	n Hα	n He	n Hβ
0.00	19.1	1.61114	1.62783	1.65238
"	28.3	.61114	.62070	.64488
13.8282	25.7	.60225	.61160	.63497
"	34.1	.59568	.60476	.62752
27.2410	24.95	.59090	.59953	.62130
"	34.1	.58404	.59259	.61410
38.605	23.7	.58077	.58898	.60953
"	34.1	.57272	.58092	.60167
44.3085	24.2	.57448	.58234	.60211
"	36.1	.56587	.57346	.59317
66.3864	26.2	.54879	.55535	.57188
"	35.5	.54196	.54852	.56515
80.7152	25.15	.53104	.53690	.55149
"	37.4	.52282	.52862	.54305
100	27.1	.50383	.50751	.51853
"	42.7	.49349	.49795	.50902

Dimethyl ether (C_2H_6O) + Sodium iodide (NaI)

Wadsworth and Dawson, 1926

%	f.t.	%	f.t.
5.6	-20	28.2	30
7.4	-10	26.3	40
10.7	0	24.2	50
15.4	+10	22.3	60
23.1	20.	20.1	70
28.6	25.7	17.9	80

Ether ($C_4H_{10}O$) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1904

%	f.t		%	sat t
0.6	- 8		40.1	23
0.8	0	(2+1)	41.8	24
0.91	+ 2		44.08	26
1.27	10		46.7	28.5
1.64	14		47.4	30
1.93	16			
2.2	17.2			
2.3	18			
2.59	19			
2.7	20			
2.87	20.3			
3.22	22			

% L_1		% L_2	sat. t
42.0	(2+1)	1.8	-10
41.0		2.3	0
40.1		2.8	+10
39.3		3.3	20
38.7		3.8	30
38.2		4.3	40
37.8		4.7	50
37.6		5.1	60
37.6		5.4	70
37.8		5.6	80
38.1		5.7	90

%	f.t.		%	f.t.
49.1	0	(1+1)	42.9	110
48.5	10		42.5	120
48.3	20		42.1	130
47.9	30		41.7	140
46.7	40		41.3	150
46.1	50		41.0	158
45.5	60		4.15	158
44.9	70		3.4	158
44.3	80		1.4	159
43.8	90		0.27	162
43.3	100		0.13	170

Menshutkin, 1904

% L1	% L2	sat. t
		(1+1)
42.0	1.8	-10
37.6	5.4	+70
38.5	5.8	100
38.9	5.7	110
39.3	5.6	120
39.7	5.4	130
40.15	5.1	140
40.6	4.7	150
41.0	4.3	158

Rowley, 1937

t	p. 0%	p. sat.sol.
0	186.7	185.7
5	234	233
10	292	290
15	361	359
20	442	438
25	537	532

mol%	p L_1+L_2	p L_1+L_2+C	p C
		25°	
27.1	529	-	-
27.9	523.5	-	-
28.7	518.5	-	-
30.1	511.5	-	-
31.6	501.5	-	-
32.8	481.5	-	-
35.1	449.5	499.5	-
37.9	-	502.5	-
37.9	-	498.5	-
38.7	397.5	492.5	-
40.3	-	483.5	-
41.0	-	498.5	-
42.3	-	494.5	-
43.1	-	480.5	-
44.6	-	489.5	-
50.0	-	404	-
50.2	-	-	389
52.1	-	-	392
53.7	-	-	385.5
54.3	-	-	389.5
78.1	-	-	392

%	f.t.	%	f.t.	%	f.t.
0.27	-20	1.44	+13.0	2.11	0
.40	-10	1.73	+15.0	2.56	+13.1
.70	0	1.95	+16.5	2.88	20.0
.83	+3.0	2.09	+18.0	3.11	24.9
1.04	+7.8	2.46	+20.0	3.22	27.5
1.17	+10.1			3.50	32.0
(3+1)		(2+1)		L_1+L_2	

Ether ($C_4H_{10}O$) + Magnesium iodide (MgI_2)

Menshutkin, 1904-05

%	f.t.	%	sat.t.
1.45	5.4 (2+1)	35.5	14.8
2.43	11.8	35.5	17.6
3.46	15.6	35.8	20
5.4	18.1	35.5	28.4
7.55	20.4	35.7	33
11.28	22.2	35.9	35
47.65	24.5	32.8	38
50.1	26	29.8	38.5
50.9	28	26.07	38.5
54.5	31.5	22.45	38.5
57.4	35.5	19.4	37.3
58.5	40.5	15.82	32.4
61.0	45	14.6	24.4
64.3	50	14.4	23.2
64.99	51	13.57	18.6

Ether ($C_4H_{10}O$) + Zinc bromide ($ZnBr_2$)

Rowley and Anderson, 1941

%	f.t.	%	f.t.
32.3	- 20	66.2	+ 10
40.2	- 15.8	69.0	+ 15
48.0	- 10	67.6	0
51.9	- 5	69.9	+ 20
57.1	0	70.2	+ 25
64.1	+ 5	70.9	+ 35

Ether ($C_4H_{10}O$) + Aluminum chloride ($AlCl_3$)

Evans, Gibb jr., and al., 1954 (fig.)

molarity	ϰ	molarity	ϰ
0.01	0.0002	0.10	0.044
.03	.004	.12	.015
.05	.010	.15	.14
.08	.025	.17	.20

Ether ($C_4H_{10}O$) + Aluminum bromide ($AlBr_3$)

Plotnikov and Kaplan, 1948

mol%	f.t.	
94.79	95.0	
87.9	82.0	
74.4	25.0	
70.9	18.0	
68.6	below -5	
64.3	" -5	
58.4	+9.0	
56.3	17.0	
50.7	42.0	
50.0	47.0	(1+1)
48.6	44.0	
45.4	22.0	

Usanovich and Terpugov, 1933

mol%	κ	mol%	κ
	18°		
8.29	0.056	38.17	2.109
10.95	.071	43.81	2.070
14.34	.309	45.85	2.2269
18.91	.477	55.43	3.080
19.02	.472	64.84	4.878
20.27	.719	72.19	5.800
21.62	.708	78.96	6.451
22.92	.934	81.43	9.269
23.65	.907	86.31	26.000
29.61	1.339	92.02	18.700
30.20	1.503	97.16	5.700
34.86	1.858	100.00	0.620

mol %	t	κ
12.41	-20.0	0.096
"	0.0	0.119
"	+18.0	0.133
"	33.0	0.145
31.60	18.1	0.725
"	20.0	0.832
"	28.3	0.961
"	33.2	1.030
"	38.5	1.110
33.35	0.0	0.574
"	10.0	0.719
"	20.0	0.805
"	30.0	1.005
"	40.0	1.160
37.41	0.0	1.560
"	10.0	1.970
"	20.0	2.480
"	30.0	2.960
"	40.0	3.480
"	50.0	3.890
57.39	0.0	1.29
"	10.0	1.98
"	20.0	2.95
"	30.0	3.34
"	40.0	5.70
"	50.0	7.20
68.46	0.0	2.35
"	24.3	6.07
"	41.0	13.80
"	60.0	20.90
86.81	22.0	28.30
"	31.7	39.80
"	40.0	44.20
"	46.0	50.20
93.02	38.5	27.30
"	50.0	34.00
"	63.5	39.00
"	70.0	42.50

Ether ($C_4H_{10}O$) + Antimony trichloride ($SbCl_3$)

Kurnakov, Perelmutter and Kanov, 1915-16 and Kurnakov, 1924

mol%	d 25°	32°	50°	75°
0	0.7076	0.6995	-	-
25	1.2172	1.2064	-	-
33.4	1.4130	1.4028	-	-
50	1.7704	1.7629	-	-
70	2.2007	2.1879	2.1512	2.1032
75	2.2959	2.2831	2.2518	2.2026
80	2.3988	2.3839	2.3418	2.2913
85	2.4900	2.4753	2.4361	2.3950
100	-	-	2.7322	2.5724

mol%	η 25°	32°	50°	75°
0	244	233	-	-
25	1008	924	-	-
33.4	1955	1678	-	-
50	6556	5269	-	-
70	15312	1573	5182	2786
75	17228	12603	5971	3051
80	18025	13349	5626	3349
85	16920	12742	6462	3429
100	-	-	4073	2438

Ethyl ether ($C_4H_{10}O$) + Germanium tetrachloride ($GeCl_4$)

Sisler, Batey and al., 1948 (fig.)

mol%	f.t. metastable	f.t. stable
0	-123.5	-115
7	-125.5 E	-117.5
8	-118.5	-118.5 E
20	-98	-
40	-80	-
60	-69	-
80	-59.5	-
100	-52	-

Ethyl ether ($C_4H_{10}O$) + Stannic bromide ($SnBr_4$)

Sisler, Schilling and Groves, 1951 (fig.)

mol%	f.t.	E
1	-20	-
10	+8	-
20	+15	-
40	+12	-
50	+9	-
52	+7	+7
60	+14	+7
80	+20	+7
100	+35	-

Ethyl sulfide ($C_4H_{10}S$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d 60	d 95	η 80	η 95
0	0.7717	-	262	-
10	1.2332	1.2042	563	477
25	.7892	.7418	2019	1161
33.3	.9952	.9436	3471	1752
36	2.0455	.9836	3932	1895
38	.1111	2.0488	4021	1983
40	.1628	.0901	3972	1993
42	.2013	.1402	3835	2022
44	.2540	.2003	3703	2001
47	.3222	.2673	3399	1994
50	.3732	.3184	3163	1869
60	.5780	.5200	2534	1638
75	.8051	.7545	1840	1396
90	3.0252	.9780	1424	1194
100	.2018	3.1516	1302	1167

Methylal ($C_3H_8O_2$) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

%	f.t.	%	f.t.
0.3 (2+1)	20	1.1	106
0.45	40	86.2	106
0.6	60	90.8	108 L_1+L_2
0.75	80	95.4	110
0.9	100	100	112

Acetal ($C_6H_{14}O_2$) + Magnesium iodide (MgI_2)

Menshutkin, 1907

%	f.t.	%	f.t.
0.15 (2+1)	20	93.7	79
0.45	60	95.5	81
0.6	77	97.3	83
92.0	77 L_1+L_2	100.0	86

Tetrahydrofuran (C_4H_8O) + Germanium tetrachloride ($GeCl_4$)

Sisler, Batey and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-108.5	60	-73.5
20	-115	80	-60.5
28	-117.5 E	100	-49.5
40	-93.5		

Tetrahydrofuran (C_4H_8O) + Stannic iodide (SnI_4)

Sisler, Schilling and Groves, 1951 (fig.)

mol%	f.t.	mol%	f.t.
0	-95	27	+80
4-5	-35	70	+120
10	+35	100	+145
20	+65		

Tetrahydropyran ($C_5H_{10}O$) + Germanium tetrachloride ($GeCl_4$)

Sisler, Batey and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-49.2	60	-75
20	-90.5	80	-62
30	-115 E	100	-52
40	-96		

Tetrahydropyran ($C_5H_{10}O$) + Stannic iodide (SnI_4)

Sisler, Schilling and Groves, 1951 (fig.)

mol%	f.t.	mol%	f.t.
0	-49	40	+95
0.9	-51.0	44	+100
10	+38	100	+145
20	+65		

Dimethylpyrone ($C_7H_8O_2$) + Aluminum bromide ($AlBr_3$)

Plotnikov, 1911 (fig.)

%	f.t.	%	f.t.
0	130	65	123
15	112	83	52
54	168	90	80
62	120	100	98

Anisole(C_7H_8O) + Antimony tricloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	tr.t.	
0	-34	-	-	
11.8	-36.5	-	-	
16.0	-30	-36.5	-	(1+1)
22.0	-20	-	-	
28.3	-10	-	-	
35.2	0	-	-	
43.0	+10	-	-	
52.8	20	-	-	
57.6	23	-	-	
63.6	25	-	-	
66.4	30	-	25	(1+2)
70.0	35	-	25	
75.9	40	-	25	
80.9	41.5	-	-	
84.5	40	-	-	
85.2	38.5	-	-	
77.3	10	-	-	(0+1)
79.3	20	-	-	
81.7	30	-	-	
84.5	40	-	-	
87.9	50	-	-	
92.0	60	-	-	
94.8	65	-	-	
98.0	70	-	-	
100.0	73	-	-	

Anisole(C_7H_8O) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	tr.t.	
0	-34	-	-	
2.5	-35	-	-	
5.4	-30	-35	-	(1+1)
11.7	-20	"	-	
18.4	-10	"	-	
26.5	0	"	-	
37.1	10	-	-	
43.5	15	-	-	
50.5	20	-	-	
59.0	25	-	-	
69.2	29	-	-	
77.0	30.5	-	-	
77.9	30	-	-	
73.2	12	-	-	(0+1)
75.4	20	-	-	
77.9	30	-	-	
80.6	40	-	30	
83.4	50	-	30	
86.4	60	-	30	
89.8	70	-	30	
93.6	80	-	30	
97.8	90	-	-	
100.0	94	-	-	

Anisole(C_7H_8O) + Germanium tetrachloride ($GeCl_4$)

Sisler and Cory, 1947

mol%	f.t.	E	mol%	f.t.	E
0.0	-37.5	-	52.0	-51.5	-
3.3	-38.0	-	55.1	-52.0	-
6.6	-39.0	-	57.6	-53.0	-
10.8	-40.5	-	63.5	-55.5	-
15.9	-42.0	-	66.9	-58.0	-
18.4	-43.0	-			
23.6	-45.5	-	61.6	-54.0	(1+2)
28.0	-48.0	-	63.5	-53.5	-
31.2	-50.0	-	66.9	-53.0	-
33.8	-52.0	-54.0	67.3	-53.0	-
37.6	-53.0	-(1+1)	70.0	-53.0	-
39.5	-52.5	-	72.3	-54.0	-
40.0	-52.5	-	76.6	-56.0	-58.0
43.5	-52.0	-	81.0	-57.0	-58.0
46.5	-51.5	-	88.0	-54.0	-
48.5	-51.0	-	93.2	-52.0	-
49.7	-51.0	-	100.0	-49.5	-
50.3	-51.0	-			

Sisler, Wilson and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-38	50	-48
10	-40.5	60	-51
20	-43	70	-54.5
30	-45.5	81	-58.5 E
40	-47	100	-52

Anisole(C_7H_8O) + Tin tetrachloride ($SnCl_4$)

Sisler and Cory, 1947

mol%	f.t.	E	
0.0	-37.5	-	
2.6	-38.5	-	
5.1	-40.0	-45.0	
7.9	-42.0	-45.0	
10.2	-44.0	-	
11.1	-45.0	-45.0	
14.2	-38.0	-	(1+1)
17.8	-34.0	-	
20.5	-32.0	-	
23.4	-29.0	-	
29.0	-26.0	-	
31.4	-25.0	-	
33.0	-24.5	-	
36.4	-23.5	-	
40.0	-23.0	-	
44.6	-22.3	-	
47.5	-22.4	-	
48.5	-22.3	-	
50.8	-22.5	-	
51.0	-22.5	-	
53.0	-22.7	-	
55.1	-23.0	-	
60.0	-22.5	-	(1+2)
61.3	-22.0	-	
63.2	-21.5	-	
65.0	-21.3	-	
68.0	-21.0	-	
70.4	-21.5	-	
72.0	-22.0	-	
74.4	-22.5	-	
77.6	-24.0	-	
79.6	-25.0	-	
81.5	-26.0	-	
83.6	-27.0	-	
85.7	-29.0	-	
87.6	-30.0	-	
89.5	-32.0	-	
91.4	-34.0	-38.5	
93.5	-38.0	-	
94.5	-36.5	-	
95.9	-35.0	-38.50	
97.2	-33.0	-	
98.9	-31.5	-	
100.0	-30.5	-	

Sisler, Wilson and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-38	60	-20
7	-40 E	66.5	-21
10	-35	80	-23
20	-27.5	90	-30.5
30	-22	94.5	-36.5 E
40	-20	100	-34
50	-19.5 (1+1)		

Usanovich and Pichugina, 1956

%	d			
	20°	40°	60°	80°
100	2.2340	2.1819	2.1288	2.0790
84.94	1.8911	1.8461	1.8069	1.7649
73.85	.6955	.6566	.6213	.5864
61.82	.5269	.4923	.4608	.4264
55.40	.4493	.4182	.3860	.3480
47.75	.3655	.3334	.3049	.2757
21.16	.1314	.1110	.0885	.0681
0	0.9978	0.9765	0.9588	0.9389

%	η			
	20°	40°	60°	80°
100	927	758	637	540
70.05	932	724	590	501
53.95	958	732	589	492
40.18	1010	746	593	491
34.01	1040	765	606	492
27.50	1070	779	626	502
10.02	1100	811	631	513
0	1080	795	627	502

Phenetole($C_8H_{10}O$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	
0	-28.6	-	
1.4	-29	-	
4.5	-20	-29	(1+1)
8.1	-10	-29	
12.3	0	-29	
18.2	+10	-29	
27.4	20	-29	
32.9	25	-	
39.7	30	-	
47.9	35	-	
58.0	40	-	
65.0	42.2	-	(1+2)
69.6	40	-	
77.8	35	-	
82.8	30	-	
85.3	26	-	
78.9	30	-	anh.
82.8	40	-	
86.8	50	-	
91.6	60	-	
94.3	65	-	
97.1	70	-	
100	73	-	

Phenetole($C_8H_{10}O$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	
0	-28.6	-	
1.6	-29	-	
2.7	-20	-29	(1+1)
4.8	-10	-29	
8.0	0	-29	
12.9	10	-26	
19.2	20	-	
29.7	30	-	
36.8	35	-	
46.2	40	-	
58.7	45	-	
66.9	47	-	
74.7	48.8	-	
77.8	47	-	
63.8	10	-	anh.
67.4	20	-	
71.2	30	-	
75.0	40	-	
79.0	50	-	
83.0	60	-	
87.3	70	-	
92.1	80	-	
97.4	90	-	
100.0	94	-	

Phenetole($C_8H_{10}O$) + Germanium terachloride ($GeCl_4$)

Sisler, Wilson and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-30	80	-53
7	-33 tr.t.	85	-55.8
20	-37	87	-57.5 E
40	-42	100	-52
60	-46		

Phenetole($C_8H_{10}O$) + Stannic chloride ($SnCl_4$)

Sisler, Wilson and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-30	62	-48
20	-38.5	80	-42
38	-49	100	-34
40	-48		

Methyl-o-cresyl ether ($C_8H_{10}O$) + Germanium tetrachloride ($GeCl_4$)

Sisler, Wilson and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-34	73	-60 E
20	-43	80	-58
40	-50	100	-52
60	-57		

Methyl-m-cresyl ether ($C_8H_{10}O$)+ Germanium tetrachloride ($GeCl_4$)

Sisler, Wilson and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-55.5	60	-64
20	-63.5	80	-59
42	-67.8 E	100	-52

Methyl-m-cresyl ether ($C_8H_{10}O$) + Stannic chloride ($SnCl_4$)

Sisler, Wilson and al., 1948

mol%	f.t.	mol%	f.t.
0	-55.5	60	-28.5
8	-59.5 E	80	-34
20	-39.5	90	-39.5 E
40	-28	100	-34
50	-27 (1+1)		

Propylphenyl ether ($C_9H_{12}O$) + Germanium tetrachloride ($GeCl_4$)

Sisler, Wilson and al., 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-28	80	-50
20	-34	88.5	-56 E
40	-38.5	100	-52
60	-43		

Diphenyl ether ($C_{12}H_{10}O$) + Germanium tetrachloride ($GeCl_4$)

Sisler and Cory, 1947

mol%	f.t.	E	
0.0	28.0	-	
3.6	26.5	-	
8.4	24.0	-	
12.4	22.0	-	
17.9	20.5	-	
21.6	18.5	-	
27.0	16.0	-	
31.5	14.0	-	
34.2	13.0	-	
37.7	12.0	-	
41.0	10.0	-	
45.0	8.7	-	(2+1)
46.3	8.0	-	
48.1	7.5	-	
49.7	7.0	-	
52.0	6.5	-	
52.5	6.0	-	
57.0	5.0	-	
57.4	4.8	-	
60.8	3.0	-	
66.8	0.5	-	
73.0	-2.0	-	
78.5	-5.0	-	
83.3	-10.0	-	
85.2	-12.5	-	
87.9	-16.0	-	
92.4	-23.0	-	
95.1	-31.0	-	
96.9	-42.5	-52.0	
98.9	-51.0	-52.0	
100.0	-49.7	-	

Diphenyl ether ($C_{12}H_{10}O$) + Stannic chloride ($SnCl_4$)

Sisler and Cory, 1947

mol%	f.t.	mol%	f.t.
0.0	28.0	42.4	13.0
1.7	27.5	44.7	12.0
2.0	27.0	46.9	11.8
4.1	27.0	50.0	10.7
5.9	26.0	52.1	10.5
9.6	25.0	57.7	8.5
13.4	23.7	65.8	5.5
17.0	22.0	70.3	4.0
18.9	21.0	72.0	3.0
20.2	20.5	76.9	0.5
21.9	20.0	81.4	-2.5
23.9	19.5	85.5	-6.5
25.0	19.5	89.6	-11.0
25.7	19.0	90.4	-13.0
28.4	18.0	91.5	-16.0
30.9	17.0	92.0	-17.5
32.3	17.0	93.9	-22.0
34.8	16.0	95.4	-30.0
37.0	15.0	96.2	-34.0 E
38.3	14.5	97.3	-33.0
39.5	14.0	98.9	-31.0
40.8	13.5	100.0	-30.5

Furfural ($C_5H_4O_2$) + Sodium iodide (NaI)

Kosakewitsch, 1928

mol%	d	σ
	15°	
0.00	1.162	43.91
1.64	.183	44.36
3.20	.205	44.83
5.75	.240	45.55

Benzaldehyde (C_7H_6O) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

%	f.t.	%	f.t.
0.7	0	7.5	125
0.9	10	9.5	130
1.1	20	12.7	135
1.3	30	17.8	140
1.5	40	22.0	142
1.7	50	30.0	144
1.9	60	37.5	145
2.2	70	65.0	146
2.5	80	75.0	147
2.9	90	84.5	148
3.4	100	89.2	150
4.2	110	93.2	153
5.0	115	97.1	156
6.0	120	100	159

Benzaldehyde (C_7H_6O) + Magnesium iodide (MgI_2)

Menshutkin, 1907

%	f.t.	%	f.t.
3.2	0	18.5	100
3.4	10	22.0	110
3.8	20	32.0	115
4.4	30	40.0	120
5.3	40	53.0	125
6.4	50	74.5	130
7.7	60	86.0	133
9.1	70	94.2	136
11.0	80	97.5	138
13.7	90	100	139
15.5	95		

Benzaldehyde (C_7H_6O) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	%	f.t.	E
38.4	-20	-	83.9	35	-
45.5	0	-	85.8	45	37.8
49.3	+10	-	88.0	55	"
51.3	20	-	90.4	65	"
59.3	30	-	93.1	75	"
64.1	35	-	96.1	85	"
70.3	40	-	98.2	90	-
77.2	41.5	- (1+1)	100.0	94	-
81.0	40	37.8			
84.4	37.8	-			

Benzaldehyde (C_7H_6O) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	%	f.t.
43.5	10	-	79.1	-5
47.5	20	-	80.2	5
52.4	30	-	81.4	15
55.8	35	-	83.0	25
60.2	40	-	85.0	35
63.1	42	-	87.5	45
68.1	43.5	- (1+1)	90.8	55
72.0	42	25	95.2	65
74.2	40	25	98.1	70
77.6	35	25	100.0	73
80.6	30	25		
83.0	25	-		

Acetone (C_3H_6O) + Lithium iodide (LiI)

P.P Kosakewitsch and M.S Kosakewitsch, 1933

mol%	d	σ
	13-14°	
0	0.798	24.14
1.56	.823	24.53
3.22	.849	24.82
6.22	.896	25.33

Acetone (C_3H_6O) + Lithium bromide (LiBr)

Bell, Rowlands and al., 1930

%	f.t.	%	f.t.
11.7	10	24.9	50
15.4	20	28.4	60
17.5	30	42.7 (2+1)	-
20.8	40		

P.P Kosakewitsch and M.S Kosakewitsch, 1933

mol%	d	σ
	13-14°	
0	0.798	24.14
1.12	.810	24.22
2.28	.824	24.56
4.37	.849	24.73

Acetone (C_3H_6O) + Sodium iodide (NaI)

Macy and Thomas, 1926

%	f.t.	d
3.2	-34	-
7.4	-12.3	0.873
11.6	0.0	.899
21.0	+15.9	.984
23.0	59.8	-
23.6	57.2	-
25.0	50.0	-
26.8	40.2	-
28.2	32.2	1.043
28.5	25.0	.062
29.2	25.7	.070
29.4	25.0	.067

(retrogr. sol.)

Bell, Rowlands and al., 1930

%	f.t.	%	f.t.
10.3	0	28.0	30
15.4	10	26.9	35
18.8	15	25.8	40
23.0	20	24.7	45
28.6	25	23.6	50
		22.6	55

(retrograd. sol.)

Evertz and Livingston, 1949

%	f.t.	%	f.t.
1.55	-76.5	5.38	-21.5
1.53	-76.5	10.83	0.5
1.77	-62.6	11.60	0.9
2.20	-67.3	23.75	20.7
2.51	-46.3		

Coleman, 1918 and Mc Bain and Coleman, 1919-20

%	d 0.0	d 25.00	d 40.00
23.11	-	0.9970	0.9766
14.85	-	.9124	.8937
9.88	-	.8666	.8481
7.092	0.8688	.8417	.8235
6.694	-	.8384	.8201
5.346	0.8543	.8273	.8091
4.026	0.8436	.8167	.7990
2.989	0.8347	.8068	.7903
1.750	0.8256	.7939	.7817
1.035	.8201	.7936	.7765
0.7894	.8183	.7918	.7748
.5789	.8167	.7902	.7731
.5591	.8166	.7900	.7730
.3175	.8148	.7882	.7712
.3076	.8148	.7882	.7712
.2868	.8146	.7880	.7710
.1949	.8140	.7874	.7705
.1534	.8137	.7872	.7702
.1305	.813	.7873	.7700
.09211	.813	.7873	.7700
.06985	.8124	.787	.7695
.0595	"	"	.7695
.03157	"	"	.7690
.02613	"	"	"
.01152	"	"	-
.01181	"	"	0.7690
.009215	"	"	"
.00467	"	"	"
.00470	"	"	"
.003763	"	"	"
.001576	"	"	"

%	κ 0.0°	κ 25.00°	κ 40.00°
23.11	-	283.7	315.6
14.85	-	256.7	275.3
9.88	-	207.9	218.6
7.092	145.1	166.7	176.2
6.694	-	161.2	169.4
5.346	118.8	137.4	144.9
4.026	98.10	113.0	117.9
2.989	52.32	60.73	63.97
1.750	52.32	60.73	63.97
1.035	34.83	40,77	43.97
0.7894	28.02	33.29	35.70
.5789	21.88	26.28	28.13
.5591	21.59	25.41	27.19
.3175	13.65	16.42	17.82
.3076	13.20	16.00	17.38
.2868	12.77	15.23	16.46
.1949	9.345	11.21	12.17
.1534	7.691	9.296	10.14
.1305	6.561	8.068	8.859
.09122	4.978	6.071	6.686
.06985	3.895	4.835	5.327
.0595	3.352	4.161	4.628
.03157	1.931	2.424	2.703
.02613	1.624	2.056	2.308
.01152	0.7650	0.9750	-
.01181	.7925	1.007	1.135
.009215	.6123	0.7834	0.884
.00467	.3222	.4130	.4822
.00470	.3327	.4253	.4695
.003763	.2630	.3371	.3833
.001576	.1109	.1440	.1624

Osaka, 1931

%	f.t.	%	f.t.
15.4 (3+1)	9.9	27.6 (0+1)	29.0
20.1	16.7	26.9	32.0
23.0	20.0	26.2	35.9
25.5	23.0	25.1	39.9
27.3	25.0	23.1	50.0
28.2	26.0		

%	p	%	p
15.4 (3+1)	109.1	27.6 (0+1)	231.4
20.1	146.1	26.9	264.3
23.0	164.7	26.2	307.5
25.5	183.8	25.1	359.5
27 3	193.8	23.1	536 0
28.2	201.7		

P.P Kosakewitsch, 1928

mol%	d	σ	mol%	d	σ
		14°			
0.00	0.798	24.13	3.55	0.870	24.77
0.93	.819	24.22	5.07	.901	25.04
2.08	.841	24.40	6.53	.920	25.12

Acetone (C_3H_6O) + Potassium iodide (KI)

Jacek, 1915

%	f.t.	%	f.t.
2.56	-84	6.24	-32.75
2.90	-84	6.54	-32.75
6.21	-76.5	5.80	-32.5
6.71	-75.5	5.30	-30
6.76	-75.	5.81	-29.5
8.83	-70.5	6.10	-27.5
15.06	-54.75	4.50	-22.5
13.50	-53.5	4.15	-18.5
11.96	-52.5	4.40	-18
12.30	-51	4.10	-18
11.90	-51	3.90	-16
10.00	-47.75	3.90	-13.5
8.61	-44.50	3.97	-13.5
8.60	-44.50	2.66	-5
7.90	-43	2.56	0
8.60	-42	2.55	0
6.54	-34	2.58	0

retrogr. sol.

Acetone (C_3H_6O) + Sodium thiocyanate (NaCNS)

Lemme, 1897

vol%	f.t.
6.10	7.65
7.75	27.20
33.85	55.15

Hughes and Mead, 1929

%	f.t.
6.41	18.8
8.68	29.2
12.34	41.9
15.69	51.0
17.63	56.0

Acetone (C_3H_6O) + Potassium thiocyanate (KCNS)

Lemme, 1897

vol%	f.t.
12.24	6.00
14.31	55.00
15.67	26.90

von Laszczynski, 1894

%	f.t.
17.18	22
16.94	58

Acetone (C_3H_6O) + Magnesium iodide (MgI_2)

Menshutkin, 1907

%	f.t.	%	f.t.
4.9	0	28.6	80
5.5	10	40.0	85
6.1	20	59.2	90
6.7	30	80.0	95
7.4	40	87.7	97.5
8.3	50	92.5	100
10.2	60	95.6	102.5
12.5	65	98.5	105
15.2	70	100	106.5
20.0	75		

Acetone (C_3H_6O) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

%	f.t.	%	f.t.
0.2	0	14.0	74
0.4	10	50.0	75
0.6	20	71.6	76
0.8	30	78.8	78
1.0	40	83.3	80
1.2	50	86.8	82
1.45	60	89.8	84
2.0	70	92.6	86
2.7	71	95.2	88
3.7	72	97.7	90
5.5	73	100	92

Acetone (C_3H_6O) + Calcium chloride ($CaCl_2$)

Bagster, 1917

t	p dissociation	
	(2+1)	(1+1)
20	41	1
26	87	-
31	139	-
36	200	-
37	-	5
42	300	-
49	478	-
50	-	15
54	623	-
57	745	-
60	840	-
62	923	-
64	983	50

Acetone (C_3H_6O) + Calcium iodide (CaI_2)

Bell, Rowlands and al., 1930

%	f.t.	%	f.t.
42.0	0	51.3	40
44.5	10	53.0	50
46.0	20	54.5	60
49.2	30	63.3 (3+1)	-

Acetone (C_3H_6O) + Calcium nitrate (CaO_6N_2)

Bell, Rowlands and al., 1930

%	f.t.	%	f.t.
17.3	0	14.7	40
14.5	10	15.5	50
14.4	20	73.2 (1+1)	60
14.6	30		

Acetone (C_3H_6O) + Zinc bromide ($ZnBr_2$)

Bell, Rowlands and al., 1930

%	f.t.	%	f.t.
78.5	20	79.3	40
78.4	30	79.2	50 (1+2)

Acetone (C_3H_6O) + Cupric chloride ($CuCl_2$)

Etard, 1894

%	f.t.
18.4	-20
18.8	+8

Acetone (C_3H_6O) + Mercuric chloride ($HgCl_2$)

Etard, 1894

%	f.t.	%	f.t.
51.4	-23	61.9	+6
52.9	-18	61.4	12
56.6	-15	61.8	15
56.7	-10	62.0	27
58.4	-8	61.9	36
59.1	-4	62.1	54
60.1	-1		

von Laszczynski, 1894

%	f.t.
49.58	0
52.59	10
55.90	18

Lemme, 1897

vol%	f.t.
71.46	10.60
86.48	55.20
92.50	30.60

Aten, 1905

mol%	f.t.	mol%	f.t.
14.5	-15.0	23.5	10.0
14.3	0.0	23.2	17.0
18.7	+10.0	22.8	25.0
	(retrogr. sol.)		

Chénevau, 1907

%	d	%	d
	18°		
36.46	1.1712	9.91	0.8753
15.31	0.9237	0	0.8000

Le Blanc and Rohland, 1896

%	d
20°	
21.05	0.9781
10.94	.8831
0	.7989

Schönrock, 1895

%	d 21.2°	d 21.7°
0	0.79472	-
36.2488	-	1.15854

Chéneveau, 1907

%	n_D
18°	
36.46	1.3996
15.31	.3735
9.91	.3682
0	.3600

Le Blanc and Rohland, 1896

%	n_D
20°	
21.05	1.3793
10.94	.3696
0	.3607

Schönrock, 1895

%	$(\alpha)_{magn.}$
16°	
0	1.0803
36.2488	0.8290

Shaw, 1913

M	κ	M	κ
	25°		
2.130	4.96	0.0164	0.163
1.050	3.62	.00840	.101
0.528	2.28	.00410	.0685
.263	1.36	.00250	.0492
.131	0.795	.00102	.0387
.0658	.445	.000508	.0303
.0339	.269		

Acetone (C_3H_6O) + Cobaltous chloride ($CoCl_2$)

Quartaroli, 1916

Specific magnetic susceptibility (see author)

Acetone (C_3H_6O) + Cobaltous bromide ($CoBr_2$)

Bell, Rowland and al., 1930

%	f.t.	%	f.t.
35.1	0	41.4	30
35.2	10	48.0	40
39.4	25	79.0 (1+1)	-
40.7	27		

Acetone (C_3H_6O) + Antimony trichloride ($SbCl_3$)

Kurnakov, Perelmutter and Kanov, 1915-16 and
Kurnakov, 1924

mol%	d			η		
	25°	50°	80°	25°	50°	80°
0	-	2.7322	-	-	4073	2158
20	-	.4603	2.4409	-	5703	2283
23	-	.4024	.3727	-	5712	2527
25	-	.3641	.3413	-	5853	2413
30	-	.3097	.2681	-	5692	2409
33.3	2.3194	.2169	.1856	18123	5309	2360
38	2.2686	-	-	16762	-	-
50	1.9575	1.9095	-	9489	4023	-
66.6	.5900	.5716	-	3546	1918	-
75	.4012	.3624	-	1078	1183	-
100	0.7866	-	-	339	-	-

Acetone (C_3H_6O) + Titanium tetrachloride ($TiCl_4$)

Wertyporoch and Altmann, 1934

M	κ	M	κ
0°			
0.0461	1.600	0.8420	10.400
.1265	2.490	1.2120	14.600
.2260	2.730	.3770	18.620
.4410	6.540	.4940	20.920
.6470	8.120	.6780	21.520

Acetone (C_3H_6O) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altmann, 1934

M	κ	M	κ
0°			
0.02550	1.734	0.07950	6.570
.03180	1.750	.09720	7.920
.04160	1.825	.11400	8.420
.04960	2.960	.14560	8.660
.06110	4.680		

Acetone (C_3H_6O) + Quinine iodobismuthate
($Bi_2C_{20}H_{26}I_8O_4N_2$)

Picon, 1934

t	gr/100cc	d	gr/100cc	d
	L_1		L_2	
9	2.64	0.8159	94.2	1.473
19	1.97	.8063	104.9	.544
29	1.54	.7981	115.3	.609
39	1.11	.7882	123.2	.673
49	0.8	.7793	130.2	.732

2-Butanone (C_4H_8O) + Sodium iodide (NaI)

Wadsworth and Dawson, 1926

%	f.t.	%	f.t.
6.8	-70	12.0	20
9.1	-60	11.3	30
12.3	-50	9.7	40
18.5	-10	8.8	50
16.7	0	6.9	60
14.8	+10	5.8	70

(retrogr. sol.)

Acetophenone (C_8H_8O) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	
0	19.5	-	
14.3	15	1	
23.1	10	1	
28.5	5	1	
31.8	1	-	
32.9	5	1	
35.4	15	1	
38.1	25	1	
41.6	35	1	
46.8	45	1	
50.6	50	1	
55.2	55	1	
62.6	60	1	
65.4	60.5	-	(1+1)
68.4	60	32	
73.7	55	32	
77.0	50	32	
79.3	45	32	
84.0	32	-	
74.0	20	-	
77.2	0	-	
81.6	20	-	
83.8	30	-	
86.2	40	-	
89.3	50	-	
93.1	60	-	
98.2	70	-	
100.0	73	-	

Acetophenone (C_8H_8O) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	
0	19.5	-	
22.7	15	1.5	
35.0	10	1.5	
43.7	5.0	1.5	
48.6	1.5	-	
52.2	10	1.5	
56.8	20	1.5	
63.3	30	1.5	
68.8	35	1.5	
75.0	37.5	-	(1+1)
79.8	35	31	
83.2	31	-	
82.2	31	-	
84.6	40	31	
86.3	50	31	
88.4	60	31	
90.9	70	31	
94.1	80	31	
98.2	90	31	
100.0	94	-	

Kurnakov, Krotkov and Oksman, 1915 and Kurnakov, 1924

mol%	d	η	d	η	d	η
	25°		50°		95°	
0	1.0224	1617	0.9981	1246	0.9891	653
25	1.6506	10606	1.6255	5623	1.5981	2478
40	2.0612	60918	2.0278	30921	.9981	5154
45	.1952	78623	.1652	38153	2.1352	6389
48	.2838	90269	.2459	39234	.2159	6698
50	.3515	96471	.2971	42482	.2587	6972
51	.3848	93579	.3243	40996	.2915	7235
52	-	-	.3685	40812	.3258	7582
54	-	-	-	-	.3818	6998
60	2.6198	53420	2.5712	22503	.5083	6024
70	.8397	31579	.7934	13529	.7693	4298
85	3.1188	-	3.0899	63208	3.0210	3982
109	-	-	-	-	3.6926	3309

Acetophenone (C_8H_8O) + Titanium tetrachloride ($TiCl_4$)

Wertyporoch and Altmann, 1934

M	κ	M	κ
	0°		
0.0461	0.2008	0.8420	0.6145
.0920	.2430	1.2120	.5560
.1820	.2710	1.4940	.4240
.2700	.2910	1.8600	.3530
.4410	.3530	2.3060	.3530
.6470	.4380		

Acetophenone (C_8H_8O) + Stannic tetrachloride ($SnCl_4$)

Wertyporsch and Altmann, 1934

M	κ	M	κ
	0°		
0.00436	0.1008	0.07950	0.3160
.00868	.1145	.09720	.2060
.01715	.1255	.11400	.1910
.02550	.1710	.14560	.1880
.04160	.2460		

Benzophenone ($C_{13}H_{10}O$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1910

%	f.t.		%	f.t.	
0	48		52.0	90	(1+1)
8.5	44		53.5	80	"
15.4	39.5	(1+1)	54.8	70	"
19.3	60	"	56.1	60	"
21.5	70	"	56.8	80	
24.0	80	"	58.0	100	
26.5	90	"	59.7	120	
29.5	100	"	61.1	130	
33.0	110	"	63.0	140	
37.0	120	"	65.5	150	
39.0	125	"	68.6	160	
42.3	130	"	72.6	170	
45.6	125	"	78.5	180	
46.8	120	"	83.3	185	
48.8	110	"	89.1	190	
50.4	100	"	93.0	192	
			100.0	194	

Benzophenone ($C_{13}H_{10}O$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1910

%	f.t.		%	f.t.	
0	48		66.0	120	(1+1)
12.0	45		67.6	110	"
19.0	42		69.0	100	"
24.7	38		70.2	90	"
28.0	50	(1+1)	71.3	80	"
30.9	60	"	72.2	70	"
33.6	70	"	73.1	60	"
36.4	80	"	74.0	50	"
39.2	90	"	75.0	38	"
42.2	100	"	78.0	50	
45.4	110	"	80.9	60	
49.0	120	"	84.2	70	
53.0	130	"	88.0	80	
58.0	140	"	90.5	85	
59.5	142	"	93.5	90	
61.5	140	"	96.4	93	
64.0	130	"	100.0	96	

Benzophenone ($C_{13}H_{10}O$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	
0	48	-	
7.7	45	35	
16.3	40	35	
21.6	35	-	
26.2	45	35	(1+1)
31.4	55	35	"
37.5	65	35	"
42.2	70	35	"
51.3	75	35	"
55.4	76	-	"
58.8	75	39	"
67.3	70	39	"
71.6	65	39	"
76.9	55	-	"
80.6	45	39	"
82.7	39	39	"
74.2 metast.	10	-	(0+1)
76.7 "	20	-	
79.7 "	30	-	
83.2	40	-	
87.0	50	-	
91.6	60	-	
97.7	70	-	
100.0	73	-	

Benzophenone ($C_{13}H_{10}O$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	
0	48	-	
11.6	45	29	
24.4	40	29	
33.5	35	29	
41.2	29	-	(1+1)
47.5	20	-	"
41.2	29	-	"
45.6	35	29	"
50.0	40	29	"
56.3	45	29	"
61.8	47.5	29	"
66.4	48.5	-	"
71.7	47.5	40	"
76.0	45	40	"
80.0	40	-	"
82.6	50	40	
85.5	60	40	
88.7	70	40	
92.4	80	40	
97.3	90	40	
100.0	94	-	

Kurnakov, Krotkov and Oksman, 1915 and Kurnakov, 1924

mol%	%	d	η	d	η
		25°		95°	
0	0	1.1064	13610	1.0496	1746
33.33	49.70	1.7223	181180	1.6349	3728
50.0	66.40	2.1185	479460	2.0088	5015
55.0	70.72	2.2525	551940	-	-
60.0	74.78	2.3923	566210	2.2694	5637
66.67	79.81	2.5929	472320	.4543	5805
70.0	82.18	-	-	.5572	5736
75.0	85.57	-	-	.7162	5654
100	100	-	-	3.6926	3309

Benzil ($C_{14}H_{10}O_2$) + Antimony trichloride ($SbCl_3$)

Vanstone, 1914

mol%	f.t.	mol%	f.t.
100	73.0	48.26	31.0
96.45	68.5	40.0	54.0
90.47	61.0	30.59	69.4
84.62	51.0	16.81	84.0
78.24	37.0	0	94.0
70.00	- (vitreous)		

Benzil ($C_{14}H_{10}O_2$) + Antimony tribromide ($SbBr_3$)

Vanstone, 1914

mol%	f.t.	E	mol%	f.t.
100	95.0	-	53.35	44.1
93.10	86.1	-	37.77	66.3
75.59	64.6	-	27.40	76.0
66.35	48.0	36.2	13.28	86.1
55.89	41.6	-	0	94.0

Acetyl chloride (C_2H_3OCl) + Gallium trichloride ($GaCl_3$)

Greenwood and Wade, 1956.

mol%	f.t.	E
100.0	77.8	-
97.9	77.6	-
91.7	76.5	-
86.1	75	-
82.0	68	-
78.8	59	-
76.8	53	-
73.5	40	-
72.0	25	0
70.6	19	0
67.4	37	-
63.8	53	-
59.6	68	-
50.0	86	-
33.3	65	-
0.0	-98	-

Benzoyl chloride (C_7H_5OCl) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1910

%	f.t.	E	
0	-0.5	-	
7.9	-4	-7.5	
12.7	-7.5	-	(1+1)
14.1	0	-7.5	"
16.3	10	"	"
18.8	20	"	"
21.6	30	"	"
25.0	40	"	"
28.8	50	"	"
33.0	60	"	"
37.5	70	"	"
42.2	80	-	"
47.1	90	-	"
48.7	93	-	"
50.6	90	-	"
52.9	80	-	"
55.1	70	-	"
57.2	60	-	"
59.1	50	-	"
61.0	40	-	"

Benzoyl chloride (C_7H_5OCl) + Gallium trichloride ($GaCl_3$)

Greenwood and Wade, 1956.

mol%	f.t.	E	
100.0	77.8	-	
94.6	77	-	
88.3	75	-	
84.7	71	-	
80.8	65	-	
77.8	57	-	
75.6	41	-	
74.0	33	-	
72.0	25	-50°	
68.7	-5	-50	
66.1	-44	-50	
63.3	-20	-	
61.1	+22	-50	
57.4	+41	-	
56.2	+42	-	
54.2	46	-	
52.6	46	-	
51.0	46.5	-	
50.3	46.5	-	(1+1)
49.6	46.5	-	
48.2	46.5	-	
46.7	46	-	
45.3	45.5	-	
43.6	45	-	
41.6	44.5	-	
39.7	44	-	
38.0	42	-	
36.9	30	-	
35.7	18	-	
34.2	15	-	
33.5	13	-	
33.1	0	-12	

Benzoyl chloride (C_7H_5OCl) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	%	f.t.	E
0	-0.5	-	58.2	15	-
17.8	-5	-23	62.9	25	-
28.7	-10	-23	68.9	35	-
36.8	-15	-23	74.9	45	-
45.0	-23	-	82.4	55	-
47.6	-15	-23	91.2	65	-
50.7	-5	-23	96.5	70	-
54.2	+5	-23	100.0	73	-

Benzoyl chloride (C_7H_5OCl) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	%	f.t.	E
0	-0.5	-	67.8	50	-6
19.5	-3	-6	74.9	60	-6
32.0	-6	-	82.1	70	-6
41.2	10	-6	89.4	80	-6
47.5	20	-6	96.8	90	-
54.0	30	-6	100.0	94	-
60.8	40	-6			

Benzoyl chloride (C_7H_5OCl) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1910

%	f.t.	E	
0	-0.5	-	
11.7	-2.5	-5	
22.2	-5	-	(1+1)
29.2	10	-	"
33.7	20	-	"
38.2	30	-	"
42.6	40	-	"
47.1	50	-	"
51.6	60	-	"
56.1	70	-	"
60.5	80	-	"
62.7	85	-	"
65.5	90	-	"
67.7	85	-	"
68.9	80	-	"
70.5	70	-	"
71.8	60	-	"
73.2	50	-	"
74.5	40	-	"
75.8	30	-	"
77.1	20	-	"
78.8	7	-	"
80.6	20	-	
82.1	30	-	
83.7	40	-	
85.6	50	-	
87.7	60	-	
90.0	70	-	
93.2	80	-	
97.2	90	-	
100.0	96	-	

Acetic anhydride ($C_4H_6O_3$) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1909

% (6+1)	f.t.	% (6+1)	f.t.
26.4	0	44.5	90
27.54	10	48.4	100
28.7	20	52.8	110
30.0	30	57.8	120
31.6	40	63.4	125
33.5	50	69.8	130
37.7	60	88.0	133
38.2	70	85.0	135
41.1	80	100	136-137

Phenylacetic anhydride ($C_{16}H_{14}O_3$) + Sodium phenylacetate ($NaC_8H_7O_2$)

Bakunin and Vitale, 1935

mol %	f.t.	E	
0.00	72	-	
18.92	64	49	
31.85	57	52	
47.40	83	54	(1+2)
58.38	104	-	
64.18	109	56	
67.79	114	-	
75.93	138	-	
82.54	157	-	
89.40	168	-	
95.01	175	-	
100.00	186	-	

Phenylacetic anhydride ($C_{16}H_{14}O_3$) + Lithium phenylacetate ($LiC_8H_7O_2$)

Bakunin and Vitale, 1935

mol %	f.t.	E
0.00	72	-
19.10	67.5	-
34.45	187	67
51.27	235	-
58.99	249	67
73.68	276	62
77.48	278	66
95.53	315	-
100.00	320	-

Methyl formate ($C_2H_4O_2$) + Stannic chloride ($SnCl_4$)

Kurnakov, Perelmutter and Kanov, 1915-16 and Kurnakov, 1924

mol%	f.t.	mol%	f.t.
10	- 5	48.9	76.2
15	26	52.3	74.2
20	50.6	53.8	73.6
25	73	56.9	72
27.5	80	58	71.2
30	82.3	59.51	70.4
32	83.1	64	68
33.3	83.3	70	65
34	83.2	74	62.4
36	82.9	78.5	59.4
38	82.5	84	54.4
40	81.7	92.3	44.2
43	80	94	38.4
47.6	76.9	100	-33
			(2+1)

Ethyl formate ($C_3H_6O_2$) + Magnesium iodide (MgI_2)

Menshutkin, 1909

% (6+1)	f.t.	% (6+1)	f.t.
15.1	0	37	45
16.2	5	44	50
17.4	10	54	55
18.8	15	68	60
20.5	20	84	65
22.5	25	93.2	68
25.0	30	96.5	69
28.0	35	100	70.5
31.8	40		

Ethyl formate ($C_3H_6O_2$) + Mercuric chloride ($HgCl_2$)

Etard, 1894

%	f.t.
29.6	-20
29.2	-3
30.0	+24
31.0	+46

Ethyl formate ($C_3H_6O_2$) + Stannic chloride ($SnCl_4$)

Kurnakov and Voskresenskaya, 1936

mol%	Dv (cc/mole)	U	Q diss.
	30°		
0	-	0.5105	-
12	-	.4611	2170
20	-	.4182	3580
25	11.3	-	-
30	12.5	.3758	5408
32	13.8	-	-
33.3	-	.3301	5824
33.6	13.7	-	-
35	-	.2804	5718
36	13.1	-	-
50	9.5	.2360	4340
70	-	.1891	2652
75	4.7	-	-
100	-	.1476	-

Kurnakov, Perelmutter and Kanov, 1915-16 and Kurnakov, 1924

mol%	f.t.	mol%	f.t.
10	- 10	33.3	49
20	+ 15	33.7	48.6
25.9	33.5	33.9	48.6
30.5	46	36	47.1
31.8	48	38.3	45.5

mol%	d		
	30°	40°	50°
0	0.9082	0.8651	0.8824
25	1.5222	1.5092	1.4958
30	1.6386	1.6232	1.6092
32	1.6839	1.6674	1.6514
33.5	1.7050	1.6885	1.6734
36	1.7331	1.7153	1.7004
50	1.8649	1.8465	1.8301
75	2.0553	2.0331	2.0115
100	2.2056	2.1816	2.1532

mol%	η		
	30°	40°	50°
0	375	342	311
25	8788	6384	4904
30	31158	19034	12463
32	52286	27615	16102
33.5	57832	29541	16758
36	44142	23514	14237
50	6871	4952	3753
75	1526	1292	1115
100	806	725	668

Ethyl formate ($C_3H_6O_2$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1936

mol%	Dv (cc/mol)	U	Q diss.
	30°		
0	-	0.523	-
10	-	.464	1040
14.3	4.7	-	-
20	-	.407	2180
25	-	.375	2790
28.9	7.5	-	-
30	-	.341	3210
33.3	-	.329	3390
33.9	7.3	-	-
35	-	.302	3340
36.8	6.8	-	-
40	-	.288	3060
43.7	5.5	-	-
50	-	.224	2520
53.7	4.9	-	-
60	-	.201	2020
70	-	.173	1420
84.3	1.7	-	-
100	-	.128	-

Kurnakov and Chernin, 1936

mol%	d				
	0°	16°	30°	45°	50°
0	0.9275	0.9200	0.9025	0.8825	0.8755
14.3	1.5971	1.5616	1.4967	-	1.4268
20.0	-	-	-	1.6824	-
28.9	2.1221	2.0667	2.0215	-	-
32.2	.2203	.1618	2.1156	-	-
33.3	.2615	.2027	-	-	2.0708
33.9	.2708	.2131	2.1641	-	-
35.0	-	-	-	2.1518	-
36.8	-	2.2856	2.2324	-	-
37.0	-	-	-	2.2231	2.2082
38.8	2.4036	2.3452	2.2925	-	-
40.0	-	-	-	2.2667	2.2544
43.7	2.4921	2.4355	2.3834	-	-
50.0	-	-	-	2.4852	2.4677
53.7	2.7117	2.6491	2.6083	-	-
60.0	-	-	-	2.6934	2.6842
80.0	-	-	-	3.0613	3.0492
84.3	-	-	3.1858	-	-
100	-	-	3.3430	3.2970	3.2800

mol%	η				
	0°	16°	30°	45°	50°
0	514	429	371	323	310
14.3	3000	2260	1290	-	852
20.0	-	-	-	1420	-
28.9	16520	5620	2650	-	-
32.2	17400	6700	2740	-	-
33.3	17450	6720	-	-	2030
33.9	17260	6720	2800	-	-
35.0	-	-	-	2350	-
36.8	-	6540	2770	-	-
37.0	-	-	-	2340	2080
38.8	15180	6400	2750	-	-
40.0	-	-	-	2340	2090
43.7	12790	5950	2720	-	-
50.0	-	-	-	2230	2100
53.7	8960	4900	2510	-	-
60.0	-	-	-	2210	2000
80.0	-	-	-	2110	1900
84.3	-	-	2260	-	-
100.0	-	-	2290	1930	1840

Ethyl formate ($C_3H_6O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956 (fig.)

mol%	f.t.	sat.t.
98	20	-
95	36	-
92	36	40
90	36	43
85	36	48
80	36 (1+4)	50
75	36	46
70	36	40
68.5	36	-
66	44	-
60	50	-
50	56 (1+1)	-
46	50	-
42	40	-

Propyl formate ($C_4H_8O_2$) + Stannic chloride ($SnCl_4$)

Kurnakov, Perelmutter and Kanov, 1915-16 and Kurnakov, 1924

mol%	d 50°	d 70°	η 50°	η 70°
0	0.8643	0.8404	396	325
25	1.3948	1.3687	5859	3239
32	.5292	.4958	14995	6066
33.4	.5644	.5200	16467	6557
35	.5743	.5414	15303	6201
50	.7328	.6971	4281	2528
75	.9546	.9102	1204	896
100	2.1520	2.0962	668	600

Butyl formate ($C_5H_{10}O_2$)+ Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956.(fig.)

mol%	f.t.	mol%	f.t.
90	0	60	20
80	8	50	30.3(1+1)
75	10	47	24
70	12	43	15

Lysenko,1956 (fig.)

mol%	d 35°	d 70°
100	1.7	1.65
75	1.6	1.52
60	1.53	1.48
50	1.61	1.59
40	1.38	1.35
30	1.23	1.22
20	1.16	1.15
10	0.99	0.98
0	0.8908 (20°)	0.85

mol%	n_D 20°	n_D 35°
100	1.61	1.59
75	1.62	1.60
60	1.62	1.60
50	1.62	1.59
40	1.57	1.56
30	1.53	1.52
20	1.48	1.47
10	1.43	1.42
0	1.3894	1.38

mol%	η 35°	η 40°	η 50°	η 60°	η 70°
80	1000	1000	500	500	500
65	8000	7000	5000	3000	1000
60	20000	15000	10000	5000	4000
50	88000	62000	32000	20000	10000
40	54000	42000	25000	17000	12000
25	7000	6000	3000	2000	1000

mol%	κ 35°	κ 40°	κ 50°	κ 60°	κ 70°
100	0	0	0	0	0
75	1.2	1.3	2.3	2.4	3.2
67	3.7	4.8	6.2	8.0	9.8
57	3.8	5.1	7.1	9.5	12.0
50	2.3	3.4	5.2	8.3	11.2
40	2.4	3.0	5.0	7.1	9.0
30	1.0	1.2	1.8	2.5	3.8
20	1.2	1.4	2.0	2.5	3.4
10	0.4	0.5	0.6	0.6	0.7

Isobutyl formate ($C_5H_{10}O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
98	39	60	82
95	53	50	87.0 (1+1)
90	63	42	73
80	72	41	68
70	78		

Isobutyl formate ($C_5H_{10}O_2$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1936

mol%	U	Q diss.	mol%	U	Q diss.
0	0.501	-	40	0.271	3005
10	.451	1115	50	.253	2524
20	.394	2210	60	.234	2027
30	.341	2960	70	.191	1538
33.3	.312	3241	100	.128	-
35	.301	3240			

Isoamyl formate ($C_6H_{12}O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956 (fig.)

mol%	f.t.	mol%	f.t.
95	9	60	43
90	17	50	51.0 (1+1)
80	29	43	42
70	38	40	30

Lysenko, 1956, (fig.)

mol%	d 50°	d 80°
100	1.68	1.62
75	1.58	1.50
60	1.48	1.41
50	1.40	1.36
40	1.30	1.25
30	1.21	1.19
20	1.08	1.03
10	0.98	0.95
0	0.84	0.82

Lysenko, 1956, (fig.)

mol%	η 50°	60°	70°	80°
78	2000	2000	1000	1000
64	4000	3000	2000	1000
50	35000	17000	11000	7000
45	42000	23000	14000	-
40	34000	21000	13500	9500
30	11000	9000	5000	4000
20	2000	1000	1000	1000
10	500	400	200	100

mol%	κ 50°	60°	70°	fig 80°
75	1	1.5	2.0	2.1
65	3.5	5.0	6.0	7.5
55	4.0	6.0	8.0	10.0
50	3.2	5.2	7.5	9.8
40	3.0	4.0	5.4	7.0
30	0.5	1.0	1.5	2.5
20	0.5	1.0	1.5	2.0
10	0	1.0	1.0	1.0

Methyl acetate ($C_3H_6O_2$) + Magnesium iodide etherate ($MgI_2.6C_3H_6O_2$)

Menshutkin, 1909

%	f.t.	%	f.t.
0.4	0	1.8	100
0.45	10	2.4	103
0.50	20	74.2	103
0.55	30	75.6	105
0.60	40	81.7	110
0.70	50	89.0	115
0.75	60	98.0	120
0.80	70	100.0	121
0.85	80		
0.90	90		

Methylacetate ($C_3H_6O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
95.0	90	70	134
90	112	55	140 (1+1)
85	120	49	142
80	128	45	141

Methyl acetate ($C_3H_6O_2$) + Stannic bromide ($SnBr_4$)

Kurnakov and Chernin, 1936

mol%	d 15°	d 25°	d 40°	d 50°
0.0	0.9407	0.9248	0.9071	0.8950
10.0	1.2706	-	-	-
15.0	-	1.5663	1.5302	-
20.0	-	-	-	1.6922
25.0	1.9492	-	-	-
28.0	-	1.9812	1.9412	-
33.0	-	2.1372	2.0912	-
38.0	-	2.2682	2.2222	-
40.0	-	-	-	2.5342
50.0	2.5982	2.5602	2.5112	-
55.0	.7322	-	-	-
60.0	.7892	-	-	-
65.0	.8982	-	-	2.7932
75.0	3.0532	3.1102	2.9642	-
85.0	-	-	-	3.0892
100	-	3.3532	3.3132	3.2995

mol%	η 15°	η 25°	η 40°	η 50°
0.0	382	378	330	300
10.0	886	-	-	-
15.0	-	893	661	-
20.0	-	-	-	715
25.0	1780	-	-	-
28.0	-	1450	967	-
33.0	-	1720	1130	-
38.0	-	1880	1220	-
40.0	-	-	-	1100
50.0	3260	2200	1480	-
55.0	3390	-	-	-
60.0	3340	-	-	-
65.0	3240	-	-	1480
75.0	3060	2410	1780	-
85.0	-	-	-	1670
100.0	-	2540	2010	1820

Ethyl acetate ($C_4H_8O_2$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggeth, Vold and Mc Bain, 1942

%	f.t.
95.5	250
97.5	280

Ethyl acetate ($C_4H_8O_2$) + Magnesium iodide etherate ($MgI_2.6C_4H_8O_2$)

Menshutkin, 1909

%	f.t.	%	f.t.
3.2	0	29.0	57.5
4.0	10	38.0	60
4.8	20	50.5	62.5
6.2	30	63.5	65
8.6	40	82.5	67.5
11.0	45	90.5	70
13.7	50	95.2	72.5
17.0	52.5	97.7	75
21.5	55	100.0	78.5

Ethyl acetate ($C_4H_8O_2$) + Mercuric chloride ($HgCl_2$)

von Laszczynski, 1894

%	f.t.	%	f.t.
22.43	0	24.17	48
22.50	13	24.67	60
23.50	35	26.46	83

Linebarger, 1894

mol%	f.t.
13.3	0
13.7	13
13.8	30
13.9	40.5
14.0	50.2

Etard, 1894

%	f.t.	%	f.t.
39.6	-50	41.6	45
40.5	-20	44	66
40.2	-14	47.8	100
40.0	-6	50.1	131
39.5	0	57.0	150
39.9	+7	59.3	180
40.2	+19		

Aten, 1905

mol%	f.t.	mol%	f.t.
9.20	-15.0	-	+17.0
9.25	0.0	9.45	+25.0
-	+10.0		

Dukelski, 1907

%	f.t.	%	f.t.
22.8	0.0	23.5	38.5
22.7	6.5	26.4	45.3
22.8	26.1		

Ethyl acetate ($C_4H_8O_2$) + Ferric chloride ($FeCl_3$)

Gladstone and Hibbert, 1897

%	n_D
	18°
27.67	50.72
18.38	50.25
11.87	50.11

Ethyl acetate ($C_4H_8O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
97.5	27	70	90
95.0	50	60	96
90	70	50	102.5 (1+1)
85	78	45	96
80	84	40	78

Lisenko, Osipov and Feodosyev, 1954

mol%	d	$\varkappa$	
	102°	97°	102°
100	1.58	0	0
80	.55	4.3	4.5
70	.52	8.4	10.5
60	.50	20	22.5
55	.49	23.4	25.6
50	.48	21.2	24.8
47	.47	19	22
43	.43	20.6	23.6
40	.40	19	20.8
30	.28	12.2	13.6
20	.15	8	8.1
0	0.80	0	0

mol%	η	
	97°	102°
100	200	0
80	1000	800
60	2500	2200
47	6400	5200
40	5400	4500
20	2800	2300
0	200	0

Lisenko, Osipov and Feodosyev, 1954 (fig.)

mol%	Q diss.	mol%	Q diss.
100	0	40	7500
80	3500	20	4200
60	7500	0	0
50	8900		

Ethyl acetate ($C_4H_8O_2$) + Stannic chloride ($SnCl_4$)

Kurnakov, Perelmutter and Kanov, 1915-16 and Kurnakov, 1924

mol%	d 25°	d 50°	d 70°
0	0.8948	0.8640	0.8353
25	1.4741	1.4398	1.4089
32.6	.5931	.5507	.5142
33.4	.6038	.5620	.5239
34.6	.6232	.5838	.5433
35.9	.6379	.5993	.5631
50	.7759	.7361	.7030
70	.9480	.9188	.8837
100	2.2132	2.1520	2.0962

mol%	η 25°	η 50°	η 70°
0	441	345	283
25	11755	4950	2751
32.6	44770	10496	4375
33.4	50318	11247	4639
34.6	46801	10934	4599
35.9	38449	10218	4387
50	6448	3233	2089
70	1893	1245	992
100	919	558	600

Trifonov, 1924

%	d	χ
	25°	
0	0.8944	-0.639
32.09	1.1953	.585
50.11	.4397	.554
57.64	.5660	.545
59.58	.6006	.542
60.98	.6148	.540
74.71	.7744	.522
100.0	2.2085	.484

Anosov, 1926

%	n_D	%	n_D
	18.5°		
0.00	1.37464	59.58	1.49777
32.09	.43387	59.97	.49801
50.11	.47653	61.66	.50007
57.39	.49307	74.71	.50215
58.77	.49587	95.07	.50958
59.19	.49619	100.00	.51390

Kurnakov and Chernin, 1936

mol%	κ 25°	κ 50°
0	0	0
5.06	0.46	0.6353
15.09	1.6907	2.7993
20.12	1.5545	3.0313
25.20	1.1347	2.6859
33.14	0.6232	1.9039
33.75	0.6989	2.5097
33.92	0.7170	2.5261
36.41	0.8240	2.6292
40.34	1.3402	2.7940
50.72	1.8405	3.4509
60.35	1.2551	1.7769
70.33	0.8279	1.0543
80.68	0.2064	0.2131
100.00	0	0

Kurnakov and Voskresenskaya, 1936

mol%	Dv (cc/mol) 25°	U	Q sol
0	-	0.428	-
15	-	.478	2621
25	10.0	.428	4509
30	-	.400	5358
32.6	14.1	.379	-
33.1	-	-	5673
33.3	14.0	.365	5728
33.5	-	.362	5786
34.6	14.1	.354	-
38	13.8	.346	5338
50	9.9	.312	4222
68	-	.251	3002
70	4.8	-	-
85	-	.190	1251
100	-	.1476	-

Ethyl acetate ($C_4H_8O_2$) + Stannic bromide ($SnBr_4$)

Kurnakov and Chernin, 1936

mol%	f.t.	mol%	f.t.
0	-83.6	31.80	-7.0
1.12	-87.0	32.00	-3.0
1.25	-88.0	32.50	-2.3
1.41	-91.5	33.30	-1.6
1.67	-87.0	34.55	-1.0
2.60	-82.0	40.00	+2.5
5.00	-75.25	41.19	3.0
10.00	-54.50	45.00	4.7
15.00	-32.50	48.50	5.5
20.00	-18.0	50.00	5.7
22.87	-13.0	55.00	5.0
25.00	-10.2	57.50	3.8
27.01	-9.0	58.50	2.8
28.10	-8.5	60.00	3.7
29.23	-8.0	63.82	7.0
30.48	-7.6	75.00	15.2
31.00	-7.2	100.00	29.9

Kurnakov and Chernin, 1936

mol%	d			
	-5°	0°	25°	50°
0	0.929	0.9233	0.8948	0.8629
15.00	-	1.4845	1.4279	1.3731
20.00	1.667	1.6548	1.5935	1.5224
25.00	-	-	1.7744	-
30.00	2.018	2.0016	1.9177	1.8428
33.00	2.094	2.0855	2.0005	1.9235
35.00	-	2.1432	2.0557	1.9771
40.00	-	2.2565	2.1519	2.0908
40.69	2.280	-	-	-
45.00	-	-	2.2937	-
50.00	2.503	2.4883	2.4107	2.3268
60.00	2.728	2.7098	2.6260	2.5557
75.00	-	-	2.9293	2.8506
100	-	-	3.351	3.2766

mol%	η						
	-5°	0°	5°	10°	20°	25°	50°
0	673	651	577	547	482	442	351
15	2900	2130	1870	1560	1250	1080	614
20	4500	2970	2700	2120	1500	1350	702
25	7510	5680	3780	2890	2000	1680	803
30	12300	7830	5470	3760	2420	1910	977
33	14100	8430	6360	4200	2700	2100	1040
35	14950	8920	6700	4550	2810	2240	1090
37	15600	-	-	-	-	-	1140
37.5	16500	9660	7080	-	-	-	-
38	17400	-	-	-	-	-	-
39	17000	10100	-	-	-	-	-
39.5	-	11000	-	-	-	-	-
40	16210	11050	7600	5100	3210	2470	1220
41	-	11100	7870	-	-	-	-
42	15900	-	8060	-	-	-	1270
42.5	-	10350	-	-	-	-	-
43	-	-	8100	-	-	-	-
44	-	-	8070	-	-	-	-
45	15450	9950	7950	5700	3590	2700	1350
46	-	-	7620	5830	-	-	-
47	-	-	-	5830	-	-	-
48	14950	-	-	-	3770	-	-
50	14250	9450	7220	5570	3880	2880	1480
51	-	-	-	-	3870	-	-
53	12760	9180	-	-	-	-	-
55	11600	8860	6750	5250	3650	2950	1520
60	10500	7570	6000	4760	3460	2800	1530
65	-	6280	5320	4370	3220	2750	1600
75	-	5250	4570	3850	3020	2720	1680
80	-	4800	-	-	-	-	-
85	-	-	3930	3450	2850	2680	1720
100	-	-	-	-	2750	2550	1770

mol%	κ		
	0°	25°	50°
0	0.0049	0.0045	0.0069
10	1.023	0.825	0.570
15	1.601	1.530	1.002
17	1.615	1.655	1.091
18	1.590	1.662	1.095
20	1.432	1.651	1.090
25	-	1.463	-
30	0.840	1.170	0.755
33	0.735	1.031	0.650
35	0.665	0.945	0.590
40	0.505	0.770	0.445
50	0.250	0.463	0.215
60	0.124	0.164	0.100
75	0.020	0.031	0.015
100	-	0.0029	0.0029

Kurnakov and Voskresenskaya, 1936

mol%	Dv (cc/mol) 25°	U	Q diss.
0	-	0.479	-
15	4.84	-	-
20	5.67	0.432	840
25	7.5	-	-
33.3	-	0.413	1310
35	7.3	-	-
40	5.4	0.387	1530
45	5.9	-	-
48	-	0.337	1680
50	7.5	0.296	1710
52	-	0.260	1690
55	-	0.210	1660
60	4.3	0.228	1530
65	-	0.183	1410
75	2.9	-	-
80	-	0.152	1010
85	-	0.149	800
100	-	0.127	-

Propyl acetate ($C_5H_{10}O_2$) + Magnesium iodide (MgI_2)

Menshutkin, 1909

% (6+1)	f.t.	% (6+1)	f.t.
4.1	0	59.5	47.5
4.74	10	72.5	50
5.4	20	81.6	52.5
6.5	30	88.2	55
6.8	32.5	92.7	57.5
7.8	35	96.0	60
10.2	37.5	98.1	62.5
19.0	40	100.0	65
32.5	42.5		
45	46.0		

Propyl acetate ($C_5H_{10}O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
97	18	60	68
95	25	55	74
90	32	50	78 (1+1)
85	45	45	71
80	50	40	54
70	58		

Osipov, Lysenko and Akopov, 1955 (fig.)

mol%	d 70°	η 70°	d 80°	η 80°
0	1.63	400	1.61	300
20	1.55	1100	1.53	1000
40	1.48	5800	1.46	4800
50	1.43	16000 (52%)	1.41	9500 (54%)
60	1.34	10800	1.31	7500
80	1.07	1300	1.06	1100
100	0.82	400	0.80	400

mol%	ϰ 70°	ϰ 80°	Q sol 23°
0	0	0	-
20	1.8	2.1	3600
40	9.2	12.2	7800
44	9.3	12.7	-
50	8.3	12.3	9400
52	7.7	11.6	-
56	8.2	11.3	-
64	4.8	6.8	6300
68	4.2	6.0	6300
75	4.6	6.0	-
80	4.2	5.3	4500
100	0.2	0.2	-

Butyl acetate ($C_6H_{12}O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956.

mol%	f.t.	mol%	f.t.
95	31	60	72
90	46	50	80.5(1+1)
80	55	45	79
70	66	43	72

Osipov, Lysenko and Akopov, 1955 (fig.)

mol%	d 80°	η 70°	η 80°	η 85°	η 90°
0	1.61	400	400	400	400
20	1.51	1400	1300	1200	1100
40	1.41	-	4800	3800	2800
50	1.36	-	8400	7000	5700
55	1.33	-	8500	7000	5700
60	1.27	10300	7200	6200	5300
80	1.02	1600	1450	1300	1200
100	0.80	400	400	400	400

mol%	ϰ 90°	ϰ 85°	ϰ 80°	ϰ 70°
0	0	0	0	0
20	3	2.5	2.2	1.6
40	12.7	11.2	9.9	-
47	12.9	11.1	9.7	-
52	13.3	11.3	9.9	-
60	9.6	8.4	7.3	6.1
67	7.9	6.9	6.1	4.7
80	4.0	3.3	2.8	2.5
90	0.7	0.5	0.4	0.3
100	0	0	0	0

Lysenko, Osipov and Feodosyev, 1954 (fig.)

mol%	Q diss.	mol%	Q diss.
20	3500	61	7800
40	7700	70	7000
49	9100	80	5000
57	7700		

Isobutyl acetate ($C_6H_{12}O_2$) + Magnesium iodide (MgI_2)

Menshutkin, 1909

% (6+1)	f.t.	% (6+1)	f.t.
10.5	0	33.7	70
12.0	10	36.5	72.5
13.6	20	40.5	75
15.4	30	45.6	77.5
17.6	40	52.0	80
20.4	50	63.5	82.5
22.3	55	89.0	85
24.9	60	95.1	86
28.7	65	100.0	87.5

Isobutyl acetate ($C_6H_{12}O_2$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1936

mol%	U	Q diss.
0	0.494	-
20	.438	810
34	.375	1303
48	.329	1576
50	.307	1610
52	.283	1640
55	.254	1603
60	.235	1582
70	.189	1364
80	.153	1085
100	.128	-

Isobutyl acetate ($C_6H_{12}O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956.

mol%	f.t.	mol%	f.t.
90	60	50	98 (1+1)
85	72	45	94
80	78	40	80
70	88	38	70
60	92		

Amyl acetate ($C_7H_{14}O_2$) + Lithium bromide (LiBr)

Sakhanov, 1910

M	λ
25°	
0.291	0.003
0.492	0.002
1.493	0.001

Amyl acetate ($C_7H_{14}O_2$) + Calcium iodide (CaI_2)

Sakhanov, 1910

M	λ	M	λ
25°			
0.0672	0.007	0.267	0.025
.110	.012	.271	.030
.138	.010	.391	.050

Amyl acetate ($C_7H_{14}O_2$) + Zinc iodide (ZnI_2)

Sakhanov, 1910

M	λ 25°
0.118	0.001
0.371	0.006
0.770	0.018

Isoamyl acetate ($C_7H_{14}O_2$) + Magnesium iodide (MgI_2)

Menshutkin, 1909

% (6+1)	f.t.	% (6+1)	f.t.
7.7	0	28.6	47.5
9.1	10	33.2	50
11.5	20	40.0	52.5
15.3	30	47.8	55
17.7	35	63.0	57.5
20.9	40	83.0	58.5
25.5	45	94.1	59
		100	60

Isoamyl acetate ($C_7H_{14}O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
80	20	50	65.3 (1+1)
75	25	45	53
70	38	42	47
60	52	40	40

Ethyl propionate ($C_5H_{10}O_2$) + Stannic chloride ($SnCl_4$)

Kurnakov and Beketov, 1915 and Kurnakov, 1924

mol%	d		η	
	25°	70°	25°	70°
0	0.8830	0.8302	529	329
10	1.0663	1.0082	1051	362
25	.3550	.2811	5997	1501
30	.4499	.3661	13710	1885
33.33	.5060	.4168	20970	2181
35	.5332	.4445	21400	2247
40	.5882	.5059	13380	2147
50	.6968	.6185	5276	2517
60	.8036	.7210	2883	1120
75	.9637	.8667	2597	807
90	2.1172	2.0077	1110	667
100	2.2132	2.0962	919	601
32.89	1.5062	1.4158	19820	2069
33.33	.5108	.4205	20120	2067
34.15	.5221	.4331	20560	2125
34.80	.5341	.4451	20520	2166

(2+1) 2nd series

Ethyl propionate ($C_5H_{10}O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
95	54	60	102
90	68	50	110.2 (1+1)
80	85	40	100
70	93	35	67

Methyl butyrate ($C_5H_{10}O_2$) + Stannic chloride ($SnCl_4$)

Kurnakov, Perelmuter and Kanov, 1915, Kanov, 1924 and Kurnakov, 1924

mol %	d		
	25°	50°	70°
33.5	1.5077	1.4572	1.4166
100	2.2132	2.1520	2.0962

mol %	η		
	25°	50°	70°
0	546	374	312
33.5	22059	4849	2092
100	919	668	600

Ethyl butyrate ($C_6H_{12}O_2$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
95	57	50	103.0 (1+1)
85	70	42	95
80	78	40	92
70	88	38	73
60	95		

Ethyl butyrate ($C_6H_{12}O_2$) + Stannic chloride ($SnCl_4$)

Kurnakov , Perelmuter and Kanov, 1915 and Kurnakov, 1924

mol%	d			η		
	25°	50°	70°	25°	50°	70°
0	0.8726	0.8468	0.8259	628	466	381
25	1.2926	1.2566	1.2247	6426	2751	1512
30.8	.3818	.3517	.3144	14612	4489	2077
32.5	.4203	.3778	.3396	18696	5032	2265
33.3	.4342	.3917	.3566	19315	5244	2332
34.0	.4435	.4000	.3634	19618	5281	2353
34.5	.4514	.4080	.3695	19828	5338	2370
35.0	.4575	.4151	.3776	19597	5363	2396
36.0	.1690	.4265	.3885	18778	5320	2406
50	.6256	.5872	.5556	4124	2864	1782
75	.9480	.9188	.8837	1893	1245	992
100	2.2132	2.1520	2.0962	919	668	600

Kurnakov and Voskresenskaya, 1936

mol%	Dv (cc/mol)	U	Q diss.
		25°	
0	-	0.473	-
10	-	0.437	1960
18	-	0.411	3470
25	11.4	0.380	4580
30.8	11.9	-	-
32	-	0.359	5350
32.5	13.2	-	-
33.3	13.4	-	-
33.4	-	0.350	5390
34	13.2	0.348	5420
34.5	13.4	-	-
35	13.3	0.340	5400
36	12.9	-	-
40	-	0.326	5310
50	9.5	0.301	4700
60	-	0.283	3940
75	6.4	-	-
80	-	0.212	2280
100	-	0.148	-

Ethyl butyrate ($C_6H_{12}O_2$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1936

mol%	U	Q diss.
0	0.463	-
15	.381	420
20	.353	514
30	.319	641
33.4	.308	668
35	.304	678
40	.284	655
50	.242	635
65	.203	521
100	.128	-

Isobutyl isobutyrate ($C_8H_{16}O_2$) + Stannic chloride ($SnCl_4$)

Bingham, 1907

C.S.T. 170°

Ethyl orthoformate ($C_7H_{16}O_2$) + Magnesium bromide-Ethyl orthoformate ($MgBr_2.2(C_7H_{16}O_2)$)

Menshutkin, 1907

%	f.t.	%	f.t.
11.1	0	25.7	80
11.7	10	29.5	85
12.5	20	35.0	90
13.5	30	41.0	95
14.8	40	50.0	100
16.5	50	66.0	105
18.6	60	88.5	110
21.5	70	100	114

Methyl carbonate ($C_3H_6O_3$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d		η	
	25°	50°	25°	50°
0	0.9234	0.7025	575	325
10	1.2345	1.0032	923	621
20	.5545	.3321	1269	861
30	.7582	.5644	1534	1069
40	2.0093	.8038	1774	1223
50	.2592	2.1038	1953	1342
60	.5823	.3992	2122	1465
75	.9456	.7123	2335	1614
100	3.3445	3.2995	2572	1823

Ethyl carbonate ($C_5H_{10}O_3$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d		η	
	25°	50°	25°	50°
0	0.9742	0.9652	778	575
25	1.4025	1.3025	1362	951
40	1.9435	1.8169	1864	1174
48	2.1384	2.0056	2252	1297
50	2.1592	2.0774	2315	1326
65	2.5083	2.4051	2521	1581
75	2.8088	2.6504	2580	1687
85	3.0235	2.8123	2610	1762
100	3.3445	3.2995	2572	1823

Ethyl carbonate ($C_5H_{10}O_3$) + Stannic chloride ($SnCl_4$)

Kurnakov and Voskresenskaya, 1937

mol%	f.t.	mol%	f.t.
0	-	38	45.9
5	15.0	39	47.5
10	30.1	40	49.0
20	42.0	45	51.5
25	45.0	46	52.1
30	47.4	48	52.8
31	47.6	49	52.9
32	47.8	50	53.1 (1+1)
33.3	47.9 (2+1)	51	52.8
34	47.5	52	52.6
35	46.5	53	52.3
36	44.1	56	51.6
37	42.1	60	50.0
		70	40.0
		75	33.5
		80	26.5
		85	16.5
		100	-30

mol%	d			
	40°	50°	65°	75°
0	0.9674	0.9652	0.9392	0.9148
10	1.0721	.9986	.9784	.9531
20	.2746	1.2242	1.1990	1.0982
30	.3898	.3374	.2850	.2064
32	.3904	.3468	.2922	.2451
33.3	.3996	.3605	.3214	.2612
34	.4211	.3803	.3429	.2812
35	.4323	.3920	.3517	.2911
37	.4492	.4066	.3640	.3001
40	.4812	.4275	.3911	.3813
45	.5149	.4632	.4213	.4021
50	.6017	.5449	.5132	.4832
60	.7562	.6852	.6235	.6033
75	.8236	.7962	.7742	.7562
100	2.1831	2.1501	2.1011	2.0840

mol%	η			
	40°	50°	65°	75°
0	626	575	442	368
10	1363	1135	934	725
20	3242	2249	1565	1143
30	7121	5132	2575	1725
32	8662	6102	-	1923
33.3	9525	6451	2851	2062
34	9754	6551	2942	2161
35	9482	6653	3051	2303
37	8151	6758	3201	2451
40	7207	6451	3301	2552
45	5920	5154	3228	2605
50	4852	3853	2825	2304
60	3171	2315	2035	1612
75	1651	1168	1125	955
100	751	651	555	451

Methyl oxalate ($C_4H_6O_4$) + Stannic chloride ($SnCl_4$)

Kurnakov and Voskresenskaya, 1937

mol%	f.t.	mol%	f.t.
0	54.2	52	89.8
5	50.8	55	88.2
10	49.2	60	88.2
15	46.2	70	88.2 L_1+L_2
20	42.5	80	88.2
25	63.2	90	88.2
30	73.2	92	88.2
40	86.8	95	72
48	90.4	100	-30
50	90.6 (1+1)		

Methyl oxalate ($C_4H_6O_4$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d		η	
	80°	90°	80°	90°
0	1.1242	1.1102	782	671
15	1.5422	1.5300	888	801
30	1.9834	1.9342	974	915
50	-	2.4092	-	1037
70	-	2.7832	-	1095
85	-	2.9552	-	1138
100	3.1960	3.1580	-	1180

mol%	f.t.	sat.t.
0	54	-
	[illegible]	-

Ethyl oxalate ($C_6H_{10}O_4$) + Stannic chloride ($SnCl_4$)

Kurnakov and Voskresenskaya, 1937

mol%	f.t.	mol%	f.t.
0	-41	60	82.3 (L_1+L_2)
10	+30.2	70	82.3
20	58.5	80	82.3
30	69.1	90	82.3
40	78.6	92	82.3
47	81.9	97	82.3
50	83.6	100	-30
52	82.6 (1+1)		

Ethyl malonate ($C_7H_{12}O_4$) + Stannic chloride ($SnCl_4$)

Kurnakov and Voskresenskaya, 1937

mol%	f.t.	mol%	f.t.
0	-50	60	110
20	-2	70	"
30	+55.8	80	"
40	98	85	" L_1+L_2
45	112.1	90	"
48	115.8	92	"
50	116.8	95	"
52	110 (1+I)	100	-30

Ethyl malonate ($C_7H_{12}O_4$) + Stannic bromide ($SnBr_4$)

Kurnakov and Chernin 1936

mol%	d 40°	d 70°	η 40°	η 70°
0.0	1.1302	1.0950	1505	947
10.0	.4019	.3014	2710	1312
15.0	.5480	.4892	3747	1659
25.0	.8002	.6792	7614	1809
30.0	.9382	.8502	9918	1935
31.0	.9592	.8712	10640	1976
32.0	.9804	.9016	11010	1961
33.3	2.0302	.9502	10410	1955
35.0	.0722	.9792	9690	1951
100.0	3.3132	3.2152	2010	1441

Kurnakov and Chernin, 1936

mol%	d 35°	d 90°	η 35°	η 90°
0.0	1.0372	0.9805	1470	680
15.0	.3842	1.2872	3950	929
25.0	.6304	.5101	9510	971
33.3	.8392	.7042	17410	1091
38.0	.9542	.8132	25370	1119
40.0	2.0122	.8632	28670	1138
42.0	.0615	.8812	34650	1136
43.0	.0812	.8942	37400	1133
44.0	.1162	.9292	37780	1131
45.0	.1306	.9451	37240	1125
50.0	.2392	2.0352	26030	1119
100.0	3.3232	3.1632	2170	1164

Ethyl ethylmalonate ($C_9H_{16}O_4$) + Stannic bromide ($SnBr_4$)

Kurnakov and Chernin, 1906

mol %	d 25°	d 40°	η 25°	η 40°
0.00	1.0001	0.9849	2109	1537
10.12	1.2018	1.1703	3301	2004
22.71	1.4372	1.4052	5314	2612
36.39	1.7252	1.5852	7065	2860
39.79	1.8062	1.7582	7549	2939
41.00	1.8206	1.7812	7651	2970
41.87	1.8522	1.8052	7720	2968
42.90	1.8652	1.8192	7670	2967
50.01	2.0206	2.0846	7064	2811
54.03	2.1262	2.1752	6218	2779
65.51	2.3709	2.3516	4617	2506
77.89	2.7122	2.6652	3476	2312
100.00	3.3532	3.3132	2536	2006

Methyl succinate ($C_6H_{10}O_4$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	f.t.	mol%	f.t.
0	18.4	50	36.8 (1+1)
2	18.0	60	36.3
5	17.0	70	35.0
7	16.5	80	32.5
9	15.6 E	85	31.1
10	21.8	88	30.0
14.5	28.0	90	27.7
20	32.2	92	25.5 E
30	35.0	95	27.0
35	35.8	100	29.5
45	36.6		

mol%	d 40°	d 60°	d 75°
0	1.1165	1.0946	1.0782
33	1.8803	1.8355	1.8123
47	2.1851	2.1487	2.1138
50	.2646	.2209	.1854
53	.3348	.2848	.2501
55	.3637	.3158	.2781
57	.4233	.3718	.3158
60	.4994	.4501	.4129
65	.5955	.5449	.5072
80	.9473	.8415	.8996
100	3.3750	3.3165	3.2720

mol%	η 40°	η 60°	η 75°
0	1787	1249	100
33	2232	1590	1256
47	2594	1649	1311
50	2599	1650	1317
53	2552	-	-
55	2534	1651	1348
57	2523	1672	1348
60	2497	1676	1350
65	2423	1664	1355
80	2242	1660	1340
100	2026	1603	1340

Ethylsuccinate ($C_8H_{14}O_4$) + Stannic chloride ($SnCl_4$)

Kurnakov and Voskresenskaya, 1937

mol%	f.t.	d	η
		100°	
0	(-20.8)	0.9224	383
10	+1.0	1.0304	545
20	-	1.1590	1025
25	37.2	-	-
30	44.6	1.2451	1718
33.3	52.1	.3040	2068
40	71.8	.3810	2860
49	-	.4825	4793
50	96.4 (1+1)	.5098	5216
52	94.2	.5252	4969
55	-	.5634	3627
60	79.7	.6062	2711
66.6	68.2	.6941	1608
75	57	.7160	783
80	49.5	-	-
90	36.3	.9618	391
100	-30.0	2.0803	242

Methyl benzoate ($C_8H_8O_2$) + Lithium bromide (LiBr)

Sakhanov, 1910

M	λ	M	λ
	25°		
0.131	0.006	1.34	0.004
0.425	0.005	1.54	0.004
0.500	0.005	2.27	0.003
0.820	0.003		

Ethyl benzoate ($C_9H_{10}O_2$) + Stannic chloride ($SnCl_4$)

Kurnakov, Perelmutter and Kanov, 1915 and Kurnakov, 1924

mol%	f.t.	mol%	f.t.
0	-	41.5	43.8
5	16	42.5	42
10	27.5	43.5	43.4
20	39.5	45	44.8
25	43.3	48	47.3
27.5	44	50	47.5
30	45	52	47.3
31.5	45.1	55	46.4
33	45.3	60	44.5
33.3	45.5	70	36.7
33.6	45.4	80	26.5
34	45.3	90	6
38	44.8	100	-33
40	44		

Kurnakov, Perelmutter and Kanov, 1915 and Kurnakov, 1924

mol%	d 25°	d 50°	d 70°	η 25°	η 50°	η 75°
0	1.0422	1.0191	1.0003	2034	1280	955
25	.3572	.3160	.2998	7896	2648	1555
33.4	.4692	.4189	.4000	11584	3200	1766
35	.4858	.4403	.4077	11779	3222	1777
37.5	.5175	.4725	.4377	11863	3305	1797
40	.5461	.5000	.4648	11848	3370	1855
45	.6058	.5622	.5243	10551	3340	1853
50	.6614	.6180	.5807	8267	3058	1743
70	.8770	.8318	.7930	2715	1558	1099
100	.2132	2.1520	.0962	919	667	600

Kurnakov and Voskresenskaya, 1936

mol%	Dv (cc/mol) 25°	U	Q diss.
0	-	0.450	-
20	-	0.400	1990
25	6.5	-	-
30	-	0.375	2700
33.4	8.0	-	-
33.5	-	0.311	2770
35	7.9	0.300	2790
36.4	-	0.298	2810
37.5	8.0	-	-
38	-	0.291	2780
40	7.9	-	-
42	-	0.235	2700
45	7.7	0.220	2620
50	7.3	-	-
55	-	0.194	2320
70	4.4	-	-
75	-	0.171	1420
100	-	0.148	-

Ethyl phenylacetate ($C_{10}H_{12}O_2$) + Sodium phenylacetate ($NaC_8H_7O_2$)

Bakunin and Vitale, 1935

mol%	f.t.	mol%	f.t.
0.00	-30	40.30	141
2.07	-41	50.34	144
6.21	-48 (1+1)	60.89	154
8.27	-31.8	70.80	160
10.34	-	80.60	167
21.60	92	90.36	173
30.79	130	100.0	186

Butyl chloracetate ($C_6H_{11}O_2Cl$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
85	38	42	60
77	55	32	45
60	67	30	35
50	72 (1+1)		

Lysenko, 1956, (fig.)

mol%	d 75°	mol%	d 75°
100	1.65	30	1.30
75	1.53	20	1.20
60	1.50	10	1.10
50	1.44	0	1.00
40	1.38		

mol%	η 75°	η 80°
100	500	500
75	1400	1100
60	2500	2200
50	4500	3300
40	4500	3800
30	3000	2500
20	2000	1800
10	1200	1000
0	700	700

mol%	κ 75°	κ 80°
100	0	0
75	0.2	0.2
65	1.0	1.2
60	1.6	2.0
50	2.2	2.5
40	2.5	2.7
30	2.0	2.4
20	0.5	0.8
10	0.1	0.1
0	0	0

Isobutyl chloracetate ($C_6H_{11}O_2Cl$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956, (fig.)

mol%	f.t.	mol%	f.t.
98	52	60	104
95	65	50	107.3 (1+1)
90	80	40	101
80	90	30	81
70	100	25	63

Isoamylchloracetate ($C_7H_{13}O_2Cl$) + + Titanium tetrachloride ($TiCl_4$)

Lysenko, Osipov and Akopov, 1956 (fig.)

mol%	f.t.	mol%	f.t.
90	30	50	74 (1+1)
80	48	40	65
70	60	35	55
60	70	30	35

Ethyl trichloracetate ($C_4H_5O_2Cl_3$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, 1956 (fig.)

mol%	d 20°	d 30°
100	1.74	1.70
80	1.65	1.63
65	1.60	1.59
50	1.55	1.54
35	1.50	1.49
20	1.44	1.43
0	1.3814	1.38

mol%	η 20°	η 25°	η 30°
100	880	850	800
80	1140	1110	950
65	1250	1200	1150
50	1450	1300	1250
35	1600	1400	1350
20	1650	1500	1370
0	1660	1520	1400

mol%	n_D 20°	mol%	n_D 20°
100	1.60	35	1.52
80	1.58	20	1.45
65	1.55	0	1.4492
50	1.54		

Butyl trichloracetate ($C_6H_9O_2Cl_3$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, 1956 (fig.)

mol%	d 20°	d 30°
100	1.74	1.70
80	1.64	1.60
65	1.55	1.53
50	1.48	1.45
35	1.43	1.40
20	1.36	1.34
0	1.2723	1.27

mol%	n_D20°	mol%	n_D20°
100	1.60	35	1.48
80	1.56	20	1.46
65	1.54	0	1.4508
50	1.48		

mol%	η 20°	η 25°	η 30°
100	880	850	800
80	1200	1100	1050
65	1550	1400	1300
50	1900	1700	1500
35	2100	1900	1700
20	2300	2100	1850
0	2350	2150	1950

Isobutyl trichloracetate ($C_6H_9O_2Cl_3$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, 1956, (fig.)

mol%	d 20°	d 30°
100	1.74	1.70
80	1.61	1.60
65	1.54	1.52
50	1.45	1.42
35	1.40	1.39
20	1.34	1.32
0	1.2618	1.25

mol%	η 20°	η 25°	η 30°
100	880	850	800
80	1280	1200	1100
65	1500	1400	1300
50	1800	1550	1450
35	2000	1800	1600
20	2100	1900	1800
0	2100	1950	1810

mol%	n_D20°	mol%	n_D20°
100	1.60	35	1.48
80	1.58	20	1.46
65	1.54	0	1.4470
50	1.52		

Isoamyl trichloracetate ($C_7H_{11}O_2Cl_3$) + Titanium tetrachloride ($TiCl_4$)

Lysenko, 1956, (fig.)

mol%	d 20°	d 30°
100	1.74	1.70
80	1.60	1.58
65	1.51	1.49
50	1.44	1.42
35	1.36	1.34
20	1.30	1.28
0	1.2300	1.22

mol%	η 20°	η 25°	η 30°
100	880	850	800
80	1250	1220	1200
65	1700	1500	1400
50	1950	1750	1600
35	2250	2000	1800
20	2500	2200	2100

mol%	n_D20°	mol%	n_D20°
100	1.60	35	1.50
80	1.58	20	1.48
65	1.55	0	1.4490
50	1.52		

Acetonitrile (C_2H_3N) + Sodium iodide (NaI)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-44.9	4.9	18.0
1.5	-50.0	4.6	50.8
4.6	-4.2	3.1	102.2
4.9	-0.1	1.8	148.5
7.5	+0.2		

(6+1)

Acetonitrile (C_2H_3N) + Potassium thiocyanate (KSCN)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-44.9	20.0	136.7
4.1	+25.0	23.7	140.4
4.6	40.0	23.9	140.6 tr.t.
4.8	43.5	27.6	143.5
5.5	65.5	39.1	148.3
6.4	79.4	47.2	151.4
7.0	87.5	58.9	154.6
9.6	112.2	65.6	156.1
13.0	124.6	100.0	178.7
17.4	134.7		

Acetonitrile (C_2H_3N) + Silver nitrate ($AgNO_3$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-44.9	42.1	25.0
9.4	-49.0 E	44.1	40.0
14.4	-25.0	47.8	69.5
22.5	-2.0	54.3	109.5
31.7	+8.6	59.8	146.0
39.4	4.5	-	159.8 tr.t.
41.2	18.0	100.0	208.6

(2+1)

Rabinovich and Sakhanov, 1915

%	d	η
	25°	
0	0.7775	35.9
4.80	.814	40.9
14.07	.893	52.4
31.91	1.092	104
56.93	1.541	648

molarity	λ	molarity	λ
	25°		
0.0461	113.0	0.929	26.00
.0313	93.6	1.686	19.30
.114	61.25	2.785	13.70
.268	44.20	3.401	11.50
.613	31.80	5.154	6.56

Acetonitrile (C_2H_3N) + Magnesium iodide (MgI_2)

Menshutkin, 1909

% (6+1)	f.t.	% (6+1)	f.t.
37.2	0	62.5	60
41.44	10	64.9	65
45.6	20	67.9	70
49.8	30	71.7	75
54.0	40	76.5	80
58.2	50	83.0	85
60.3	55	91.3	89

Acetonitrile (C_2H_3N) + Zinc iodide (ZnI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-44.9	2.9	88.1
0.8	+25.0	4.2	104.9
1.0	40.0	5.2	115.3
1.7	61.3	6.5	126.2

Acetonitrile (C_2H_3N) + Cadmium iodide (CdI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-44.9	4.3	40.0
-	-47.0 E	3.7	78.8
6.7	-13.0	3.4	93.9
4.9	+25.0	3.1	110.3

Acetonitrile (C_2H_3N) + Mercuric chloride ($HgCl_2$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-44.9	14.0	76.0
8.7	+25.0	16.0	88.5
8.8	27.0	18.6	104.7
10.0	40.0	25.1	132.6
10.3	41.0	37.4	167.7
11.9	57.8	100.0	276.0
13.3	69.0		

Acetonitrile (C_2H_3N) + Cobalt nitrate (CoN_2O_6)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-44.9	15.6	61.0
0.5	+31.4	17.2	63.9
1.1	37.2	22.7	74.1
1.8	42.7	27.2	85.5
3.3	44.7	32.5	89.6
5.3	45.7	40.5	90.8
6.7	45.7	51.7	87.0
8.5	48.2	53.7	84.4
10.0	50.8	59.0	79.5 E
12.3	50.7	66.9	109.0

(6+1),(4+1),(3+1),(3+2)

Acetonitrile (C_2H_3N) + Titanium chloride ($TiCl_4$)

Wertyporoch and Altmann, 1934

M	κ	M	κ
	0°		
0.1265	3.420	1.2120	9.920
.5720	6.225	.3350	10.290
.8420	7.150	.4900	11.130
1.0250	9.160	.8600	12.000

Acetonitrile (C_2H_3N) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altmann, 1934

M	κ	M	κ
	0°		
0.0436	9.950	0.4160	72.000
.0868	16.500	0.7950	122.000
.2130	40.100	1.1400	117.000
.3360	59.200	1.4560	91.200

Propionitrile (C_3H_5N) + Titanium tetrachloride ($TiCl_4$)

Hertel and Demmer, 1932 (fig.)

mol%	f.t.	mol%	f.t.
24	107.5	37.7	106.5 E
25	108.0 (3+1)	38.0	109
25.7	106.0 E	40	118.5 (3+2)
26	107.0	41.3	110 E
28	114.5	42	112.5
28.25	114.0 (5+2)	44	118.5
30	109.5	45.6	120 (6+5)
31	105.5 E	46	120
32	108.5	48	110
33.2	112 (2+1)	49.2	104 E
34	111	50.0	106.0 (1+1)
36	108.5		

Puschin, Ristic and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	-23	-	50	100	(1+1)
93	80	-	48.5	99	
88	85.5	-	47.5	98	
82	-	-23.5	45	93	
81	88	-	44	91.5	
76	88.5	-	34.5	74.5	
70	89.5	-	26	56	
68	90	-	17.3	12.5	
64.5	91.5	-	17	8.5	
61	93.5	-	0	-104	
54.5	97.5	-			

Propionitrile (C_3H_5N) + Stannic chloride ($SnCl_4$)

Pushin, Ristic and al., 1942

mol%	f.t.	E	mol%	f.t.
100	-33	-	35	75
89	+23	-	33.4	76.5 (2+1)
76.5	35	-34	30	73.5
65.5	45	-	28	71.5
56	54	-35	26	66
44	66	-	17	31
40	71	-	0	-104

Acrylonitrile (C_3H_3N) + Aluminum chloride ($AlCl_3$)

Grebenyuk and Tsukervanik, 1955 (fig.)

mol%	f.t.	mol%	f.t.
80	10	50	118 (1+1)
70	175	48	100
67	181 (1+2)	40	110
55	95	38	147

Succinonitrile ($C_4H_4N_2$) + Silver nitrate ($AgNO_3$)

Middelberg, 1903

%	f.t.	%	f.t.
0	55.6	29.6	22.7
4.8	46.92	29.7	23.55
10.0	38.4	33.7	27.67
13.5	32.9	35.0	31.2
15.45	29.5	38.4	37.7
18.1	24.4	42.3	44.2
19.75	20.7	47.7	54.4
22.7	15.5	51.2	59.8
24.85	16.95 (1+1)	52.9	62.75
26.0	18.3	55.0	66.65
26.65	19.55	61.0	74.4
26.8	19.3	65.5	78.25
29.0	22.8	68.4	85.8

Benzonitrile (C_7H_5N) + Silver nitrate ($AgNO_3$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.	
0.0	-13.8	47.5	130.5	
8.7	-17.8 E	51.8	149.3	
19.8	+ 9.1	54.1	159.8	tr.t.
32.3	13.1	58.0	169.8	
38.9	18.0	69.5	189.2	
40.0	76.1	100.0	208.6	
42.7	98.3			

(2+1)

Benzonitrile (C_7H_5N) + Mercuric chloride ($HgCl_2$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-13.8	9.3	59.3
3.8	+25.0	10.8	71.6
4.8	29.5	12.6	87.5
6.0	37.8	14.8	103.1
6.1	40.0	19.7	130.3
7.3	42.9	28.4	163.5
8.5	47.2	100.0	276.0

(2+1)

Benzonitrile (C_7H_5N) + Aluminum bromide ($AlBr_3$)

Müller, 1932

%	f.t.	%	f.t.
1st series		2nd series	
22.97	- 8	23.48	- 5
23.51	- 5	23.94	0
24.21	- 2	22.86	+ 5
23.95	0	21.09	15
22.85	+ 5	27.46	25
22.35	10	29.62	40
21.09	15	30.26	60
23.25	18	31.35	80
25.12	22	32.27	95
27.48	25		
29.05	30		
29.62	40		
30.17	50		
30.24	60		
30.53	70		
31.35	80		
31.84	90		
32.24	95		
32.56	100		

Benzonitrile (C_7H_5N) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	
0	-13.2	-	
10.2	-16	-	
17.2	-19	-	
24.2	-25	-	
21.9	-19	-19	(1+1)
28.5	0	"	
32.9	5	"	
38.7	10	"	
47.4	15	"	
62.6	20	-	
68.7	21.5	-	
72.4	20	-	
78.9	15	-	
81.6	25	15	
84.5	35	15	
87.6	45	15	
91.3	55	-	
95.6	65	-	
98.3	70	-	
100	78	-	

Benzonitrile (C_7H_5N) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	%	f.t.
0	-13.2	82.5	35
19.2	-16	75.8	-5
28.7	-18	77.1	5
40.2	-22	78.6	15
48.3	-27	80.4	25
		82.5	35
28.7	-18 (1+1)	84.9	45
35.2	-10	87.5	55
43.0	0	90.3	65
51.0	+10	93.3	75
59.0	20	96.5	85
67.0	30	98.3	90
71.7	35	100.0	94
77.3	38		

Benzonitrile (C_7H_5N) + Antimony triiodide (SbI_3)

Pushin, Ristic and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	171	-	38.5	144.5	-
87	165	-	30	141.5	-
74.5	160	-	20	135	-16
66	157	-26	10	117	-
57.5	154	-24	6	92	-15
46	149	-	0	-12.5	-

Benzonitrile (C_7H_5N) + Titanium tetrachloride ($TiCl_4$)

Hertel and Demmer, 1932 (fig.)

mol%	f.t.	mol%	f.t.
30.2	172	40	178
32	178	42	175.5
33.2	178.5 (2+1)	42.5	175
34	178	44	181
36	176	45.2	184 (6+5)
36.8	175	46	183.8
38	178.5	48	170.05
39.95	178 (3+2)	50	178 (1+1)

Pushin, Ristic and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	-23	-	40	179	-
89	137	-	33	180	-
75	157	-25	23	166	-
64.5	166	-	13.5	115	-20
50	174	-	0	-12.5	-
45.5	177	-			

(2+1)

Wertyporoch and Altmann, 1934

M	κ	M	κ
	0°		
0.0461	0.3727	0.4410	0.4320
.0920	.4380	.6470	.4350
.1820	.4160	.8420	.4360
.2700	.4350		

Benzonitrile (C_7H_5N) + Stannic chloride ($SnCl_4$)

Pushin, Ristic and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	-33	-	35	108	-
87	54.5	-	33.4	109	-
80	66	-34	23.5	104	-
69	81	-	20	100	-
60	91.5	-41	11	80	-
53.5	98	-	0	-12.5	-
46.5	104	-			

(2+1)

o-Methylbenzonitrile (C_8H_7N) + Stannic chloride ($SnCl_4$)

Pushin, Ristic and al., 1942

mol %	f.t.	E
100	-33	-
89.5	-36	-
80	-47.5	-35
69.5	-56	-
59	-63	-
49.5	-67	-
40	-70	-
33.3	-73	-
30.5	-71.5	-
20.5	-54	-
16	-41	-19
0	-14	-

(2+1)

o-Methylbenzonitrile (C_8H_7N) + Stannic bromide ($SnBr_4$)

Pushin, Ristic and al., 1942

mol %	f.t.	E
100	30	-
89.5	27	25
79.5	37	-
70	44.5	-
59	51	-
54.5	52	-
50	53	-
47	52.5	-
40	52	-
33.4	51	-
30.5	48	-
21.5	38	20
14	25	-
0	-14	-

(1+1)

m-Methylbenzonitrile (C_8H_7N) + Stannic chloride ($SnCl_4$)

Pushin, Ristic and al., 1942

mol%	f.t.	mol%	f.t.
100	-33	40	94.5
91.5	14	33.4	97
82	42	30	95
70	62	20	86
60	75.5	11.5	64.5
50	86.5		

(2+1)

p-Methylbenzonitrile (C_8H_7N) + Antimony trichloride ($SbCl_3$)

Pushin, Ristic and al., 1942

mol %	f.t.	E
100	73	-
90	62	-
80	48.5	8
70	30	9
61	19	9
60	21.5	-
55	24	-
50	32	-
45	28	-
44	25.5	-
40	21.5	-
37	20	-
30	16	-
25	10	-
20	13	0
10	25.5	-3
0	30	-
	(1+1)	

p-Methylbenzonitrile (C_8H_7N) + Titanium tetrachloride ($TiCl_4$)

Pushin, Ristic and al., 1942

mol %	f.t.	E
100	-23	-
90	92.5	-
80	107	-
70	121.5	-24.5
60	133	-
50	142	-
40	151	-
34	153	-
30	140	-
25	122	9"
19.5	99	-
13.5	69	25
0	30	-
	(2+1)	

Azobenzene ($C_{12}H_{10}N_2$) + Antimony trichloride ($SbCl_3$)

Vanstone, 1914

%	mol%	f.t.		%	mol%	f.t.	
100	100	73.0		74.08	69.67	76.1	
95.94	95	65.6		65.09	59.97	65.4	
93.47	92	71.8		55.60	50.16	59.2	
91.74	89.93	74.2		45.43	40.09	50.5	E
87.88	85	79.5		34.71	29.93	54.8	
86.40	83.63	80.0		23.66	19.94	59.2	
82.27	80	80.7	(1+4)	12.08	9.98	63.1	
78.88	75.05	78.3		0	0	68.0	

Azobenzene ($C_{12}H_{10}N_2$) + Antimony tribromide ($SbBr_3$)

Vanstone, 1914

%	mol%	f.t.	E
100	100	95.0	-
95.28	91.08	85.6	-
88.77	80	73.4	72.4 (1+4)
85.57	75	72.2	-
82.18	70	71.2	-
74.78	60	66.7	-
66.41	50	60.2	-
58.55	42.69	52.9	-
49.09	32.79	55.4	52.8
26.85	15.66	61.9	-
13.62	7.39	64.2	-
0	0	68.0	-

Azobenzene ($C_{12}H_{10}N_2$) + Stannic chloride ($SnCl_4$)

Pushin, 1948

mol %	f.t.	E
100	-33	-
90.5	+13.5	-33
80	23.5	-
60.5	40	-
51	46	-
40.5	52	-35
34.5	54.5	-
26	53	-
12	63.5	-
5.5	66	-
0	68	-
	(2+1)	

Methylamine (CH_5N) + Lithium chloride (LiCl)

Bonnefoi, 1901

t	p dissoc. (1+1)	t	p dissoc. (2+1)	t	p dissoc. (3+1)
339.2	505	300	149	295.4	289
344	642	313	347	308	600
347.4	819	323.2	642	313.5	827

Fitzgerald, 1912

N	d	η
	0°	
2.005	0.7802	934.6
1.351	0.7530	547.7
0.9842	0.7365	429.8
0.5266	0.7060	316.6
0.3546	0.7060	280.6

N	λ	N	λ
	0°		
2.028	10.109	0.174	7.090
1.099	12.968	.0944	5.527
0.595	11.129	.0512	5.027
0.321	8.482		

Methylamine (CH_5N) + Lithium nitrate ($LiNO_3$)

Fitzgerald, 1912

N	λ	N	λ
	0°		
3.60	10.9	0.855	21.9
1.93	16.6	0.820	22.1
1.50	21.0	0.459	15.7

Methylamine (CH_5N) + Sodium nitrate ($NaNO_3$)

Fitzgerald, 1912

N	d	η	N	d	η
		0°			
1.076	0.7711	374.2	0.0813	0.6969	243.9
0.5750	.7300	296.2	.0452	.6939	242.8
0.3866	.7139	273.9	.00570	.6908	236.8
0.2251	.6960	256.8	.00305	.6884	238.0
0.1207	.7036	248.5	.00265	.6873	236.0

Methylamine (CH_5N) + Potassium iodide (KI)

Fitzgerald, 1912

N	d	η	N	d	η
		0°			
1.02150	0.8670	439.1	0.04143	0.6940	241.3
0.06493	.7865	325.3	.02215	.6939	239.6
0.04372	.7537	291.2	.01442	.6917	237.3
0.03004	.7335	275.3	.005086	.6778	236.4
0.01471	.7109	255.3	.002711	.6889	237.5
0.07868	.6995	246.7	.001830	.6876	237.3
0.05296	.6947	242.7			

N	λ	N	λ
	0°		
1.70	41.6	0.500	36.8
1.09	43.5	0.448	35.4
0.909	42.9	0.219	24.6
0.833	43.2	0.118	18.2
0.600	37.4	0.0637	14.4

N	λ -33.5°	λ -15°	λ 0°	λ +15°
1.641	33.08	40.57	45.60	49.42
0.840	35.05	40.95	44.37	46.63
0.431	30.28	33.43	36.06	35.49
0.221	23.05	24.33	24.33	24.37
0.1131	18.50	19.09	18.35	16.79
0.0580	16.01	15.92	14.79	13.36
0.0297	15.66	15.07	13.86	12.16
0.0152	16.90	16.32	14.75	12.67

Methylamine (CH_5N) + Silver iodide (AgI)

Fitzgerald, 1912

N	d	η	N	d	η
		0°			
2.311	1.170	507.6	0.346	0.7672	268.4
1.235	0.9464	353.6	.123	.7134	249.6
1.082	0.9274	340.2	.0716	.7071	242.7
0.5788	.8188	282.7	.0383	.6992	239.9
0.3899	.7729	268.6	.0258	.6939	240.3

Next Page

Methylamine (CH_5N) + Silver nitrate ($AgNO_3$)

Fitzgerald, 1912

N	d	η	N	d	η
		0°			
3.571	1.261	3357	0.330	0.7511	298.7
2.764	1.137	1797	.176	.7206	267.8
1.995	1.042	900.4	.119	.7101	257.3
1.477	0.9354	655.3	.0991	.7074	251.9
1.285	0.9010	553.1	.0530	.6985	243.9
1.042	.8658	487.8	.0451	.7013	242.9
0.9950	.8543	451.2	.0357	.6934	241.4
0.5731	.7839	341.6	.0162	.6937	239.2
0.3751	.7534	303.6	.00865	.6939	238.0
			.00311	.6908	236.1

Franklin and Gibbs, 1907

N	λ	N	λ
	20.3°		
4.098	8.14	0.205	26.9
2.762	17.2	.189	26.4
1.721	26.4	.103	22.8
1.64	27.0	.0952	22.1
1.52	25.9	.0515	20.4
1.15	30.4	.0474	19.6
0.820	33.3	.0258	17.8
.781	31.4	.0238	19.4
.758	35.3	.0129	22.2
.410	31.5	.006453	26.2
.373	30.8	.001618	44.8

Fitzgerald, 1912

N	λ			
	-33.5°	-15°	0°	+15°
4.072	3.441	6.569	9.846	13.87
3.563	5.257	9.22	13.14	17.56
2.088	15.23	21.86	27.29	32.74
1.827	17.15	23.57	29.24	34.32
1.070	24.03	30.74	36.28	41.21
0.9354	24.46	31.23	36.28	40.99
0.4750	25.85	31.34	35.05	37.15
0.1835	23.17	26.77	28.50	29.14
0.09407	21.30	23.59	24.17	23.59
0.04822	20.28	21.73	21.73	20.78
0.02470	20.79	21.50	21.29	19.57
0.01678	21.39	22.43	21.23	20.08
0.008606	25.47	26.63	25.36	22.90
0.007158	26.94	28.03	26.88	24.47
0.004407	30.54	31.42	30.42	27.75
0.003667	32.59	33.64	32.20	29.36
0.001879	40.35	42.00	40.35	36.51
0.0009634	49.83	52.69	51.25	46.92

Methylamine (CH_5N) + Zinc chloride ($ZnCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (5+1)
37.5	62
51.5	99
65.5	260
75.5	440

Methylamine (CH_5N) + Zinc bromide ($ZnBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (5+1)
45	94
57	242
65	384

Methylamine (CH_5N) + Zinc iodide (ZnI_2)

Ephraim and Linn, 1913

t	p dissoc. (5+1)
41	196
50	299
57	390

Methylamine (CH_5N) + Cadmium bromide ($CdBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (3+1)
43	136
55	268
69.5	565
77.5	815
83	975

Methylamine (CH_5N) + Cadmium iodide (CdI_2)

Ephraim and Linn, 1913

t	p dissoc. (3+1)	t	p dissoc. (3+1)
43.5	48	83.5	330
54.5	95	98.5	365
67	170	108.5	765

Methylamine (CH_5N) + Mercuric iodide (HgI_2)

Fitzgerald, 1912

molarity	λ	molarity	λ
	0°		
1.25	3.20	0.3508	1.47
0.9560	3.00	.2392	1.08
.6506	2.46	.1865	0.91
.4504	1.83	.1269	0.71

Methylamine (CH_5N) + Cupric chloride ($CuCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
50.5	65	99	490
64.5	107	108.5	675
82.5	233	112.5	780
91	345		

Methylamine (CH_5N) + Cupric bromide ($CuBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
64.5	60	106	470
76.5	145	111.5	590
89.5	251	119.5	775
99	362	123	865

Methylamine (CH_5N) + Cupric iodide (CuI_2)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
57	70	112.5	487
67.5	105	119	620
76.5	140	124.5	760
88.5	203	129.5	898
103	345		

Methylamine (CH_5N) + Manganese chloride ($MnCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (x+1)	t	p dissoc (y+1)
41.5	170	96.5	143	143.5	65
57	362	106	272	157.5	110
64.5	520	115	393	165	150
72.5	726	122	500	175	200
79.5	945	129.5	620	185	250
84.5	1093	136	725		

Methylamine (CH_5N) + Manganese bromide ($MnBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (x+1)
67.5	95	109	38
76.5	150	119.5	76
89.5	274	133.5	153
101	460	149	291
108	595	159	400
112.5	711	171	550
117	845	179.5	670
122	995		

Methylamine (CH_5N) + Manganese iodide (MnI_2)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (x+1)
70	58	114.5	115
85	170	129.5	257
96	265	143	468
105	387	154.5	665
113.5	540	161	769
122	750		
127	900		
132	1050		

Methylamine (CH_5N) + Ferrous chloride ($FeCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (5+1)	t	p dissoc. (5+1)
59	127	90	604
71.5	278	98.5	802
80.5	424	105	992

Methylamine (CH_5N) + Ferrous bromide ($FeBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (5+1)	t	p dissoc. (5+1)
85	76	117.5	478
92.5	110	125.5	670
101	199	130.5	826
110	330	136	998

Methylamine (CH_5N) + Ferrous iodide (FeI_2)

Ephraim and Linn, 1913

t	p dissoc (5+1)	t	p dissoc. (5+1)
87.5	75	125	526
103	175	131	665
112.5	289	136	802
120.5	430	142	976

Methylamine (CH_5N) + Cobalt chloride ($CoCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (x+1)
60	150	104	56
72.5	265	114.5	83
82.5	395	128.5	141
90	505	133.5	168
97.5	625	150.5	272
103	750		
108	890		

Methylamine (CH_5N) + Cobalt bromide ($CoBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (6+1)
74.5	64	114.5	373
82.5	83	125.5	624
93	126	131.5	786
103.5	212	137.5	964

Methylamine (CH_5N) + Cobalt iodide (CoI_2)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (x+1)
95	71	104	56
104	97	114.5	83
116.5	157	128.5	141
121	195	133.5	168
126.5	251	150.5	272
135	365	160	339
140	434		
145.5	522		
150.5	605		
158	755		
163	893		

Methylamine (CH_5N) + Nickel chloride ($NiCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (6+1)
83.5	42.5	125	431
93.5	80.5	134	611
103	139	144	866
115	261		

Methylamine (CH_5N) + Nickel bromide ($NiBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (6+1)
110.5	49	162.5	398
122	66	170	554
132.5	111	172.5	602
145.5	196	180.5	803
155	292		

Methylamine (CH_5N) + Nickel iodide (NiI_2)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (6+1)
143.5	64	196	576
159	129	206	331
168	186	212	1008
184.5	366		

Ethylamine (C_2H_7N) + Lithium chloride (LiCl)

Bonnefoi, 1901

t	p dissoc. (1+1)	t	p dissoc. (2+1)	t	p dissoc. (3+1)
362.3	757	342.4	743	330	696
364.7	858	345	840	334	843

Fitzgerald, 1912

M	λ -33.5°	λ -15°	λ 0°	λ +15°
2.3724	-	-	1.686	2.211
1.2159	-	2.127	2.601	2.829
0.6234	1.359	1.874	2.031	1.951
0.3193	0.9019	1.0540	1.037	0.8559

Ethylamine (C_2H_7N) + Silver iodide (AgI)

Elsey, 1920

N	λ	η	N	λ	η
		-33.5°			
2.156	0.001913	-	0.3246	0.006216	-
1.116	.003600	996.5	.2253	.00739	-
1.068	.003721	-	.1728	.008318	637.3
1.023	.003331	-	.1012	.01002	604.6
0.6641	.004841	-	.0878	.01093	597.7
0.6339	.004707	799.4	.05629	.01349	595.1
0.5514	.005085	-	.02888	.01820	589.4
0.3494	.006337	692.2	0.01460	0.02562	585.6
			0.00000	-	574.9

Ethylamine (C_2H_7N) + Silver nitrate ($AgNO_3$)

Elsey, 1920

N	η	λ
	- 33.5°	
1.999	7098	3.180
1.004	-	5.633
0.7411	1374	5.762
0.5420	970	5.587
0.5184	-	5.204 ?
0.4917	-	5.310
0.2785	806.9	4.231
0.2731	-	3.982
0.2508	-	3.820
0.1797	-	3.263
0.1402	710.1	2.715
0.1358	-	2.517
0.1220	-	2.283
0.08535	-	1.951
0.07045	627.1	1.690
0.06791	-	1.522
0.06665	-	1.700 ?
0.04400	-	1.280
0.03436	594.8	1.191
0.03311	-	1.102
0.02559	-	1.107
0.02201	-	0.9874
0.01719	584.9	0.9944
0.01673	-	0.8937
0.01260	-	0.9945
0.01086	-	0.9284 ?
0.00850	576.4	0.9814
0.008196	-	0.9525
0.006240	-	1.052
0.005428	-	1.007
0.004277	575.7	1.008
0.004140	-	1.090
0.003117	-	1.254
0.002170	575.7	1.336
0.001485	-	1.651
0.0009003	-	2.196
0.0007315	-	2.208
0.0004566	582.3	3.003
0.0003533	-	2.864 ?
0.0002293	575.2	4.132
0.0001771	-	3.960
0.0001016	573.7	5.618
0.00006145	575.4	7.621

Fitzgerald, 1912

M	ϰ			
	-33.5	-15	0	+15
2.498	2.269	4.240	6.191	8.581
1.255	5.645	8.241	10.726	12.873
1.007	6.03	8.972	11.21	13.31
0.5047	6.187	8.105	9.654	10.90
0.252	4.592	5.740	6.528	7.142
0.126	2.852	3.381	3.672	3.922
0.0635	1.783	1.9325	2.061	2.061
0.0318	1.288	1.357	1.366	1.263
0.01596	1.103	1.121	1.074	0.9626
0.00800	1.106	1.106	1.027	0.9073

Ethylamine (C_2H_7N) + Manganese chloride ($MnCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
-17	17	+4	378
-10	54	10.5	594
-6	107	16	776
-0.5	251	22	1026

Ethylamine (C_2H_7N) + Manganese bromide ($MnBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
11.0	94	28.5	512
20.5	224	32.5	776
25.0	352	36.5	1005

Ethylamine (C_2H_7N) + Cobalt chloride ($CoCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
12	110	36.5	700
21	205	38	825
30	408	39.5	950
34	550		

Ethylamine (C_2H_7N) + Manganese iodide (MnI_2)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (4+1)
-32	100	12.5	101
-20	302	19.5	157
-15.5	397	26	226
-10	520	31	298
-5	656	40	496
-4	696	45	694
0	838	49.5	927
+4	1003	51	1006

Ethylamine (C_2H_7N) + Cobalt bromide ($CoBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
30.5	93	62	631
42	194	65.5	752
52.5	371	70	927
57.5	506		

Ethylamine (C_2H_7N) + Cobalt iodide (CoI_2)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (6+1)
40.5	52	-17	100
51	120	-10	179
58.5	202	-5	292
65.5	304	0	420
76	510	+3.5	500
82	656	7.5	600
87	803	11	700
91.5	949	13.5	800
		18.5	970

Ethylamine (C_2H_7N) + Nickel chloride ($NiCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
49	36	102	414
60.5	51	108	548
71	105	112.5	698
87.5	232	117.5	940

Ethylamine (C_2H_7N) + Nickel bromide ($NiBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (4+1)
20	34	103.5	44
37.5	81	113.5	110
42	108	124.5	191
53	200	138.5	318
62.5	322	150	538
72.5	517	159.5	775
79	690	164	951
84	843		
90	1095		

Ethylamine (C_2H_7N) + Nickel iodide (NiI_2)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (4+1)	t	p dissoc. (x+1)
57	48	118.5	33	143	44
72	154	137.5	82	158	58
84	327	145	135	176	73
90.5	455	155.5	254	204	93
95.5	558	164	403		
100.5	697	171	584		
107.5	969	177.5	802		
181.5	904				

Propylamine (C_3H_9N) + Nickel chloride ($NiCl_2$)

Ephraim and Linn, 1913

t	p dissoc. (2+1)	t	p dissoc. (1,5+1)
54.5	16	43	27
68.5	48	60	52
82.5	99	81	113
101	157	95	167
119	221	110	235
134.5	297	122.5	297
145.5	407		
154.5	511		

Butylamine ($C_4H_{11}N$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggeth, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
0.2	30	85.5	200
0.4	50	91	250
1.0	90	95	280
58.0	150		

Dimethylamine (C_2H_7N) + Silver iodide (AgI)

Elsey, 1920

N	η	λ
	-33.5°	
2.695	-	0.0000590
2.03	-	.0065030
1.366	799.2	.0002217
1.259	-	.0024310
0.7011	-	.0004584
.6375	555.8	.0004819
.3195	502.3	.0007660
.1592	461.2	.0011100
.1285	-	.0065030
.06389	441.5	.0089740
.03197	439.4	.0110900
.01629	436.8	.1438000
.008316	436.0	.1919000

Dimethylamine (C_2H_7N) + Nickel bromide ($NiBr_2$)

Ephraim and Linn, 1913

t	p dissoc. (4+1)	t	p dissoc. (4+1)
22	19	56	383
36	81	64.5	550
39	116	76	835
45.5	201	81	1015

Dimethylamine (C_2H_7N) + Nickel iodide (NiI_2)

Ephraim and Linn, 1913

t	p dissoc. (6+1)	t	p dissoc. (4+1)
0.5	41	43	90
3	101	56	155
9.5	246	72.5	290
12	310	81	377
19.5	537	90.5	533
23.5	702	100	826
26	816	104.5	1014

Cyclohexylamine ($C_6H_{13}N$) + Sodium iodide (NaI)

Lafontaine and Philippe, 1958

mol%	f.t.	mol%	f.t.
0.0	-17.4	12.5	65.5
10.1	+20.0	10.6	81.0
10.6	23.0	10.1	86.2
12.5	34.5	8.7	100.5
14.3	40.0	6.1	133.0

(3+1)

Cyclohexylamine ($C_6H_{13}N$) + Cuprous iodide (CuI)

Lafontaine and Philippe, 1958

mol%	f.t.	mol%	f.t.
0.0	-17.4	20.3	75.0
0.5	+2.0	21.7	75.6
0.8	19.4	23.9	78.3
1.4	25.2	24.6	78.5
2.2	31.5	25.0	80.0
2.3	33.9	28.7	85.5
3.4	38.9	32:1	85.7
4.6	43.5	32.7	88.9
5.9	48.6	34.5	103.2
8.2	52.3	34.9	105.2
9.4	52.4	39.1	120.9
13.4	65.1	42.4	133.5
16.3	69.5	48.1	140.5
18.1	72.8		

(6+1),(3+1),(2+1),(1+1)

Cyclohexylamine ($C_6H_{13}N$) + Silver nitrate ($AgNO_3$)

Lafontaine and Philippe, 1958

mol%	f.t.	mol%	f.t.
0.0	-17.4	7.9	86.3
0.4	+43.5	9.5	90.5
0.6	52.2	11.0	91.0
0.8	62.4	12.7	93.5
1.5	63.5	13.9	97.8
2.0	66.5	14.3	99.9
3.0	70.1	15.7	103.7
3.9	74.3	18.2	110.0
4.5	75.9	20.5	114.5
6.0	80.8	24.1	121.0
7.1	82.3	27.7	126
7.5	83.3		

(3+1)

Cyclohexylamine ($C_6H_{13}N$) + Thallium nitrate ($TlNO_3$)

Lafontaine and Philippe, 1958

mol%	f.t.	mol%	f.t.
0.0	-17.4	7.2	65.0
7.2	+21.5	4.0	75.3 tr.t.
8.3	25.0	1.9	92.0
10.6	33.5	0.9	123.0
13.9	41.0	0.5	144.6 tr.t.
13.9	49.0	100.0	206.0
10.6	56.8		
(2+1)	(retrogr. sol.)		

Cyclohexylamine ($C_6H_{13}N$) + Cadmium iodide (CdI_2)

Lafontaine and Philippe, 1958

mol%	f.t.	mol%	f.t.
0.0	-17.4	19.2	132.1
10.6	32.5	19.5	135.1
13.1	93.0	20.6	141.6
14.2	105.8	23.5	151.3
15.3	115.0	24.8	155.8
16.9	124.0	26.2	161.8
18.2	129.3	27.0	165.4
18.7	130.6		
(3+1) (2+1)			

Cyclohexylamine ($C_6H_{13}N$) + Mercuric iodide (HgI_2)

Lafontaine and Philippe, 1958

mol%	f.t.	mol%	f.t.
0.0	-17.4	9.5	39.0
2.2	-19.8	10.1	48.7
6.2	-21.9	10.4	52.2
7.7	-22.4	11.7	63.0
7.9	-22.6 E	14.9	86.8
8.7	+3.5	17.7	102.6
9.5	39.0	24.1	120
(2+1)		100.0	250

Cyclohexylamine ($C_6H_{13}N$) + Nickel acetate ($NiC_4H_6O_4$)

Lafontaine and Philippe, 1958

mol%	f.t.	mol%	f.t.
0.0	-17.4	4.0	139.5
0.6	+25.0	4.2	141.5
1.1	30.1	5.2	143.0
1.7	37.1	5.3	147.0
2.8	45.2	6.5	150.0
3.8	51.3	9.0	157.0
4.0	53.8 metast.	12.6	162.0
4.2	53.8	16.3	163.7
5.2	55.3	16.9	164.7
5.3	58.5	18.3	168.8
6.5	57.8	19.7	171.8
(6+1) (4+1)			

Aniline (C_6H_7N) + Cuprous iodide (CuI)

Lafontaine and Philippe, 1958

mol%	f.t.	mol%	f.t.
0.0	-6.1	10.8	128.0
1.4	+25.0	11.2	129.0
2.2	40.0	11.8	129.3
5.8	109.0	11.2	138.0
7.8	124.0	10.8	153.0
9.4	126.9	(6+1)	

Aniline (C_6H_7N) + Silver nitrate ($AgNO_3$)

Sakhanov, 1913

N	λ	N	λ
	25°		
1.841	0.62	0.0545	0.39
1.754	0.76	0.0379	0.35
1.097	1.57	0.0306	0.34
0.641	1.96	0.0187	0.32
0.308	1.54	0.00885	0.33
0.147	0.85	0.00535	0.37
0.0960	0.64		

Aniline (C_6H_7N) + Silver perchlorate ($AgClO_4$)

Hill and Macy, 1924

%	f.t.	dt?	%	f.t.	d
0	-6.15	-	30.9	55.0	1.281
0.74	-6.6	1.030	33	48.3	-
5.00	25.0	1.063	35	56.1	1.401
11.40	40.0	-	38.9	66.6	-
17.00	50.1	-	40	88	-
17.06	60.52	-	100	480	-

Aniline (C_6H_7N) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

%	f.t.	%	f.t.	%	f.t.
(6+1)		(4+1)		(2+1)	
3.2	10	24.0	103	76.3	237
3.5	20	24.1	110	76.6	240
3.9	30	24.3	120	77.3	250
4.4	40	24.6	130	78.1	260
5.1	50	24.9	140	79.0	270
6.0	60	25.4	150		
7.5	70	26.0	160		
9.7	80	26.8	170		
12.8	90	28.3	180		
15.1	95	30.3	190		
18.5	100	33.5	200		
22.0	102	38.5	210		
27.5	103.5	45.0	220		
		48.0	225		
		55.0	230		
		65.0	234		
		76.3	237		

Aniline (C_6H_7N) + Magnesium iodide ($MgC_{36}H_{42}N_6I_2$)

Menshutkin, 1907

%	f.t.	%	f.t.
3.3	0	11.6	140
3.5	20	14.0	145
3.7	40	17.5	150
3.9	60	22.1	155
4.1	70	27.5	160
4.3	80	32.0	170
4.6	90	52.0	180
5.0	100	64.5	188
5.8	110	65.9	200
7.0	120	67.2	210
8.5	130	68.5	220
9.8	135	69.8	230

Aniline (C_6H_7N) + Zinc bromide ($ZnBr_2$)

Menshutkin and Butkov, 1926

mol%	f.t.	mol%	f.t.
0.40	76.5	31.9	270
0.98	105	33.3	270
1.23	145	36.5	269
1.84	152	39.8	263
2.42	163	40.0	262
3.3	172	42.0	262
8.9	205	46.3	242
9.7	207	48.0	236
15.5	228	54.0	215
16.65	231	64.8	172 E
17.4	232.5	73.7	285
19.9	240	84.5	325
23.8	250	100	349
27.2	259		
27.9	260		

Aniline (C_6H_7N) + Zinc iodide (ZnI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-6.1	2.7	119.6
0.7	+25.0	4.5	143.8
1.5	91.1		

Tombeck, 1900

t	p dissoc. (2+1)	t	p dissoc. (2+1)
5	2.5	83	340.0
20	12.5	91	475.0
47	67.5	100	645.0
63	147.5	111	887.5
75	245.0		

Aniline (C_6H_7N) + Cadmium bromide ($CdBr_2$)

Tombeck, 1900

t	p dissoc. (2+1)	t	p dissoc. (2+1)
19	15	81	355.0
31	32.5	89	432.5
42	55.0	93	570.0
56	117.5	100	750.0
67	195.0	103	835.0
75	275.0		

Aniline (C_6H_7N) + Cadmiunm iodide (CdI_2)

Tombeck, 1900

t	p dissoc. (2+1)	t	p sat.sol.
20	20	132	182.5
45	55	155	335.0
53	70.0	170	445
63	95.0	183	565
70	120.0	184	570.0
81	155.0	190	650.0
100	222.5	197	735
105	235.0	202	815
125	305.0		
142	385		
155	442.5		
166	510		
173	555		
177	580		
183	625 (f.t.)		

Aniline (C_6H_7N) + Mercuric bromide ($HgBr_2$)

Staronka, 1910

%	mol%	f.t.	
15.43	4.5	62.5	(2+1)
22.30	6.9	75	
29.37	9.7	84	
39.05	14.2	93.5	
47.27	18.8	100	
55.14	24.1	106	
61.83	29.5	109	
63.07	30.6	109.5	
64.77	32.3	110.5	
66.50	33.9	111	
67.87	35.3	112	(1+1)
69.73	37.3	116	
76.37	45.5	122	
79.41	49.9	124	
82.79	55.4	123	
96.21	39.6	117.7	

Aniline (C_6H_7N) + Mercuric iodide (HgI_2)

Staronka, 1910

%	mol%	f.t.	
23.43	5.9	12	(2+1)
30.37	8.2	22.5	
35.92	10.3	29	
46.09	14.9	41.5	
49.28	16.6	45	
54.81	19.9	48.5	
62.93	25.8	53.5	
66.92	29.3	105	
69.38	31.7	122	
70.06	32.4	53	tr.t.
70.62	33	128	
72.96	35.6	140	
74.55	37.5	147	
75.89	39.2	156	

Pearce, 1914

%	f.t.	
0	-8.0	
-	-11.48	E
18.93	-6.5	(2+1)
22.29	+0.4	
30.00	17.8	
32.22	21.1	
35.68	26.9	
38.29	30.1	
43.12	36.2	
49.11	42.9	
-	46.8	t.tr. - HgI_2 yellow
56.16	48.8	
62.09	63.6	
64.80	70.8	
66.85	76.2	
71.15	95.9	
-	108	t.tr. - HgI_2 red
73.81	115.7	
74.04	137.2	
74.87	181.1	
89.64	199.1	

Pearce, 1915

N	λ		
	0°	25°	35°
0.1	0.0033	0.0076	0.0097
0.5	.0014	.0036	.0046
1	.0012	.0032	.0046

Aniline (C_6H_7N) + Mercuric cyanide ($Hg(CN)_2$)

Staronka, 1910

%	mol%	f.t.		
		stable	unstable	
9.45	3.7	41	-	(4+1)
12.27	4.9	-	26	
14.10	5.7	49	30.5	
18.47	7.7	58.5	35	
21.57	9.2	65	38.5	
31.00	14.2	77	-	(2+1)
31.70	14.6	79	-	
37.65	18.2	83.5	-	
39.98	19.7	84	-	
45.33	23.4	88.5		

Aniline (C_6H_7N) + Antimony trichloride ($SbCl_3$)

Kurnakov, Krotkov and Oksman, 1915 and Kurnakov, 1924

mol%	%	d	η	d	η
		95°		125°	
0	0	0.9554	724	0.9284	231
25.0	44.76	1.4227	6059	1.3875	1082
33.33	54.87	.5801	15360	-	-
45.0	66.55	.7997	34550	-	-
50.0	70.86	.8877	37853	1.8421	3564
52.5	72.88	.9342	37444	.8866	3571
55.0	74.82	.9305	35440	.9309	3544
66.67	82.94	2.1718	17520	-	-
100	100	.6286	1532	2.5536	497

Efremov, 1941 (fig.), (data of Menshutkin)

mol%	f.t.	mol%	f.t.
0	-8	40	90 E
10	+67	50	100.5 (1+1)
16	78	60	85
24	88 (3+1)	75	31 E
28	86 E	90	61
33	94.5 (2+1)	100	72

also: (4+1) 80° and (5+1) incongruent.

Aniline (C_6H_7N) + Antimony tribromide ($SbBr_3$)

Kurnakov, Voskresenskaya and Gurovich, 1938

mol%	f.t.	d 50°	d 65°	d 85°
0	-	0.9932	0.9807	0.9676
10	-	1.3161	1.3008	1.2810
25	43.0	1.7885	1.7744	1.7602
28	55.5	-	-	-
33.3	78.1	2.0607	2.0465	2.0210
40	95.4	.2632	.2418	.2217
45	100.5	.4265	.4078	.3807
47	102.0	.4843	.4612	.4001
49	102.2	.5410	.5203	.4931
50	103.0	.5840	.5632	.5292
51	102.8	.6001	.5812	.5539
53	102.7	.6532	.6339	.6124
55	-	.7133	.6951	.6671
60	98.2	.8420	.8261	.8032
65	90.7	-	-	-
70	80.1	-	-	-
75	61.6	3.1948	3.1751	3.1472
82	68.5	-	-	-
90	84.6	-	-	3.5041
100	94.0	-	-	-
		(1+1)		

mol%	f.t.	η 50°	η 65°	η 85°
0	-	1805.8	1327.4	936.5
10	-	6997.2	4211.5	2356.3
25	43.0	8236	49720	14204
28	55.5	-	-	-
33.3	78.1	2175200	283450	44010
40	95.4	7012200	749360	84873
45	100.5	20677000	1292000	115630
47	102.0	23031000	1397300	123180
49	102.2	24418000	1437500	126270
50	103.0	24690000	1496900	127730
51	102.8	20783000	1363000	125020
53	102.7	16036000	1155900	118150
55	-	7887500	876540	106980
60	98.2	3022400	353480	65122
65	90.7	-	-	-
70	80.1	-	-	-
75	61.6	272810	81615	25598
82	68.5	-	-	-
90	84.6	-	-	7608
100	94.0	-	-	-

Aniline (C_6H_7N) + Trimethyltin iodide (SnC_3H_9I)

Kraus and Greer, 1923

mol%	f.t.	mol%	f.t.	
93.12	52	33.35	95.2	
79.39	77	27.78	94.5	
65.79	88	23.8	93	
49.23	93.5	21.04	92	(2+1)
39.16	95	18.17	90.7	

o-Toluidine (C_7H_9N) + Zinc bromide ($ZnBr_2$)

Tombeck, 1900

t	p dissoc. (2+1)	t	p dissoc. (2+1)
15	5.0	100	340.0
43	27.5	105	382.5
61	65.0	112	483.5
80	165.0	120	592.5
88	225.0	128	722.5
95	277.5	135	840.0

Methylaniline (C_7H_9N) + Antimony tribromide ($SbBr_3$)

Kurnakov, Voskresenskaya and Gurovich, 1938

mol%	d		η	
	50°	65°	50°	65°
0	0.9643	0.9479	1401	911
25	1.6672	1.6409	21025	10180
33.3	1.9115	1.8766	107710	29200
40	2.1119	2.0701	506930	86716
45	.2589	.2270	1423200	171030
50	.4328	.3899	2866900	290800
53	.5199	.4822	3472300	342610
55	.5488	.5214	3692300	357100
57	.6314	.5910	3331000	343700
60	.7088	.6646	2685500	324230
65	.8660	.8118	1274800	219830

Benzylaniline ($C_{13}H_{13}N$) + Antimony trichloride ($SbCl_3$)

Vanstone, 1914 and 1925 (fig.)

mol%	f.t.	mol%	f.t.
100	73	40	98
90	50	30	82
85	36 E	20	60
80	56	13	32 E
70	87	10	33
60	99	0	35
50	103 (1+1)		

Benzylaniline ($C_{13}H_{13}N$) + Antimony tribromide ($SbBr_3$)

Vanstone, 1914 and 1925 (fig.)

mol%	f.t.	mol%	f.t.
100	96	40	81
90	71	30	63
80	61	20	43
76	30 E	17	28.5 E
70	50	10	31
60	72	0	35
50	85 (1+1)		

Phenylhydrazine ($C_6H_8N_2$) + Magnesium bromide ($MgC_{36}H_{48}N_{12}Br_2$)

Menshutkin, 1907

t	%	t	%
20	3.0	99	54.3
30	4.74	100	54.8
40	7.0	120	57.8
50	10.5	140	60.8
60	16.4	160	64.4
70	23.5	180	68.4
80	33.0	200	73.4
90	44.5		

p-Bromaniline (C_6H_6NBr) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

mol%	f.t.	E	mol%	f.t.
0	63.5	-	41.6	90
2.1	62	-	46.0	73
5.4	81	62	49.3	63
8.9	91	"	52.0	57
11.5	95.5	"	54.4	49
18.7	103	"	78.6	46 E
23.5	105	-	88.6	63
27.3	105	-	94.7	70
32.9	102	- (3+1)	100	73
38.2	96	-		

p-Bromaniline (C_6H_6NBr) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

mol%	f.t.	mol%	f.t.
100	94	16.1	52
92.0	92	11.5	57
86.0	83	4.8	61.5
77.5	69	1.4	62.5
69.0	35	0	63.5
22.1	40		

Kurnakov, Voskresenskaya and Gurovich, 1938

mol%	f.t.	mol%	f.t.
0	63.8	50	85.1 (1+1)
10	60.4	52	84.4
15	56.1	55	82.6
20	47.5	60	78.3
25	-	65	71.3
33.3	74.1	70	39.1
40	82.5	75	58.2
45	84.1	85	81.5
48	84.8	100	94.0

mol%	d			
	50°	65°	95°	85°
0	-	1.5296	1.4989	1.5090
25	2.1380	2.1128	2.0616	2.0840
33.3	.3392	.3002	.2523	.2731
40	.5031	.4648	.5307	.4332
45	.6236	.5848	.5307	.5514
48	.7012	.6614	.6052	.6245
50	.7505	.7115	.6593	.6796
52	.8084	.7622	.6909	.7210
55	.8738	.8332	.7712	.7928
60	.9916	.9502	.8890	.9062
75	3.3086	3.2694	3.2106	3.2316
100	-	-	3.6912	-

mol%	η			
	50°	65°	85°	95°
0	-	2279	1594	1371
25	212700	53980	14518	8935
33.3	2115000	237030	37020	19174
40	12175000	691550	68486	31859
45	20001000	1015800	94545	40190
48	20528000	1060700	95192	42081
50	20729000	1064500	95288	43131
52	19953000	953200	95029	42952
55	11637000	830800	86750	40630
60	4143200	445650	6869	34788
75	239830	73161	23005	14831
100	-	-	-	3505

1-Naphtylamine ($C_{10}H_9N$) + Zinc chloride ($ZnCl_2$)

Menshutkin and Butkov, 1926

mol%	f.t.	mol%	f.t.
0	48	20.25	251
0.16	47.9	29.5	248
3.1	197	40	220 E
8.9	220	43.8	225
15.7	245	65.5	248
20	253 (4+1)	100	262

1-Naphtylamine ($C_{10}H_9N$) + Zinc bromide ($ZnBr_2$)

Menshutkin and Butkov, 1926

mol%	f.t.	mol%	f.t.
0	48	21.8	234
0.28	47.2	24.7	230
2.28	166	25	225 E
4.6	192	32.0	242
12.65	218	43.5	253
13.7	233	55.2	275
20.0	241 (4+1)	100	349
21.3	235		

2-Naphtylamine ($C_{10}H_9N$) + Zinc chloride ($ZnCl_2$)

Menshutkin and Butkov, 1926

mol%	f.t.	mol%	f.t.
0	111	38.5	265
6.24	210	44.6	262
13.1	235	54.9	250
22.35	250	60.5	220
23.1	252	63.8	205 E
26.85	262	70.0	235
31.4	269	82.5	255
33.3	275 (2+1)	100	262

2-Naphtylamine ($C_{10}H_9N$) + Zinc bromide ($ZnBr_2$)

Menshutkin and Butkov, 1926

mol%	f.t.	mol%	f.t.
0	111	29.4	262
0.4	170	30.0	261
0.8	189	33.3	261 (2+1)
5.6	226	33.8	262
11.4	245	34.8	260
13.4	248	43.5	234
16.1	250	46.0	215
18.8	252	50.0	200 E
20.0	254	59.6	245
23.2	257	69.4	317
25.7	259	100	349

Piperidine ($C_5H_{11}N$) + Cuprous iodide (CuI)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-10.5	21.0	112.8
12.6	+49.3	22.4	124.3
14.2	68.9	23.6	135.6
16.2	84.5	26.9	157.7
19.5	103.8		
(2+1)			

Piperidine ($C_5H_{11}N$) + Cadmium iodide (CdI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-10.5	16.8	103.6
9.0	+25.0	23.4	141.5
9.8	40.0	29.5	172.4
13.0	73.6		

Pyrrole (C_4H_5N) + Thallium nitrate ($TlNO_3$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-18.4	14.3	51.7
-	-24.1 E	16.1	52.8
0.6	+6.0	17.4	54.9
1.3	14.2	18.8	71.6
2.7	21.7	20.2	89.7
4.1	29.0	21.4	101.7
5.9	33.2	24.3	122.5
7.7	38.2	26.9	135
9.8	43.5	31.4	149
11.1	45.7	38.6	165
		100.0	206.0
(3+1)			

Pyridine (C_5H_5N) + Lithium chloride (LiCl)

Kahlenberg and Krauskopf, 1908

%	f.t.	%	f.t.
11.31(2+1)	8	11.60(1+1)	40
11.31	20	11.43	49
11.37	28	11.38	60
		11.71	80
		13.01	100

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	20°	
0	0.977	36.07
2.33	0.989	35.32
6.19	1.004	36.69

Pyridine (C_5H_5N) + Lithium bromide (LiBr)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	14°	
0.983	0.983	37.15
1.000	1.000	37.50
1.021	1.021	38.09

Pyridine (C_5H_5N) + Lithium iodide (LiI)

P.P. and M.Sm Kosakewitsch, 1933

mol%	d	σ
	14°	
0	0.983	37.15
2.33	1.011	37.75
5.12	1.047	38.56
8.35	1.089	39.68

Pyridine (C_5H_5N) + Sodium iodide (NaI)

P.P. and M.S.Kosakewitsch, 1933

mol%	d	σ
	14°	
0	0.983	37.15
1.69	1.010	37.59
2.63	1.024	37.71
4.39	1.054	38.21

Pyridine (C_5H_5N) + Potassium thiocyanate (KSCN)

Wagner and Zerner, 1910

%	f.t.	
100	173.8	
92.7	172.9	$L_1 + L_2$
43.1	172.7	
3.1	43.3	E
2.2	+10.0	
1.23	70.71	
0.39	116.117	
0.4		$L_1 + L_2$ higher than 200°
3.1	-43.3	
2.4	-42.8	pyridine
1.33	-42.4	
0.5	-42.2	
0	-42.0	

Pyridine (C_5H_5N) + Cuprous iodide (CuI)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-41.5	15.1	77.8
7.3	+25.0	21.6	97.8
8.3	40.0	28.3	108.3
9.8	50.1	30.0	128.9
12.0	63.4	31.8	144.1

(2+1)

Pyridine (C_5H_5N) + Silver nitrate ($AgNO_3$)

Walden and Centnerszwer, 1906

%	D.b.t.	%	D.b.t.
2.26	+0.330	1.62	+0.278
4.06	0.550	4.24	0.558
8.25	1.082	6.71	0.869
10.95	1.480	10.03	1.344
12.96	1.831	12.62	1.771
16.43	2.508	15.02	2.229

Kahlenberg and Brewer, 1908

%	f.t.		%	f.t.	
0	-48.5		15.83	-22.0	(3+1)
2.91	-50.5		16.69	-10.0	
5.66	-53.0		18.48	0.0	
8.26	-59.0		21.39	+10.0	
-	-65.0	E	25.17	20.0	
9.99	-51.25		29.01	30	
10.47	-44.0		34.86	40	
11.12	-35.0		38.37	45	
12.20	-30.0	(6+1)	38.68	46	
14.96	-25.0		39.88	47	
-	-24.0	tr.t.	41.47	48	
			-	48.5	tr.t.
41.13	45.0	(2+1)	69.27	225.4	nitrate
41.95	50.0		69.74	230.4	
44.01	60		70.34	237.1	
47.12	70		70.76	241.9	
54.79	80		71.73	253.8	
63.25	87		73.07	271.4	
69.97	80				
69.75	74				

Müller, 1932

f.t.	mol%	f.t.	mol%
100	42.271(2+1)	25	16.284
85	41.933	19	15.888
62	42.432	11	15.867
50	42.531	5	13.706
45	42.381	3	12.601
44	42.232	0	11.536 (6+1)
40	34.793(3+1)	-10	11.575
35	24.693	-25	11.523
30	16.268	-30	11.542

Sakhanov and Przheborovski, 1915

%	d	η
	25°	
0	0.9773	891
4.044	1.018	1104
9.59	1.055	1345
14.88	1.129	2050
26.07	1.280	5670

N	λ	N	λ
	25°		
1.748	8.89	0.0647	28.7
1.058	15.62	.0424	30.6
0.800	18.18	.0175	33.8
0.490	22.3	.0112	36.6
0.251	25.2	.00569	41.6
0.157	26.7		

Pyridine (C_5H_5N) + Silver perchlorate ($AgClO_4$)

Macy, 1925

%	f.t.	%	f.t.
0	-40.3	34.5	60.4
4.0	-41.5	41.1	66.8
7.2 E	-43	41.7	68.0
8.3(4+1)	-35	42.4	71.0
12.23	-11.5	43.0	75.0
14.53	-1.3	45.8	86.3
20.90	+25.0	47.8	95.6
24.52	36.1	50.0(2+1)	110
29.4	49.2		

Pyridine (C_5H_5N) + Zinc chloride ($ZnCl_2$)

Mason and Mathews, 1925

%	f.t.	%	f.t.
1.60	0	4.66	45
1.80	5	5.85	55
2.02	10	7.33	65
2.28	15	9.06	75
2.55	20	11.11	85
2.80	24.12	13.60	95
3.67	35	16.25	105

Pyridine (C_6H_5N) + Zinc iodide (ZnI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-41.5	11.0	100.5
3.1	+18.0	12.4	110.1
3.9	37.8	13.1	114.9
5.1	50.1	14.5	121.9
6.0	59.8	15.5	128.3
6.7	67.0	17.7	141.0
7.6	74.5	20.2	141.0
9.2	88.0		

Pyridine (C_6H_5N) + Cadmium iodide (CdI_2)

Tombeck, 1900

t	p dissoc. (2+1)	t	p dissoc. (2+1)
5	2.5	61	152.5
23	20.0	79	327.5
33	37.5	90	507.5
53	102.5	100	752.5

Muchin, 1913

%	f.t.	%	f.t.
0.69	61.3	5.74	81.5
0.79	63.5	5.93	82.1
1.90	70.9	8.29	84.9
1.97	71.4	11.50	87.9
2.16	72.2	12.97	89.5
3.39	75.6	13.4	89.7
5.31	80.9	18.05	93.6

Pyridine (C_5H_5N) + Mercuric chloride ($HgCl_2$)

Walden and Centnerszwer, 1906

%	D b.t.	%	D b.t.
4.58	+0.486	4.52	+0.492
8.69	0.990	8.52	0.970
12.05	1.426	12.66	1.567
15.68	2.032	15.83	2.076
18.70	2.580	18.70	2.607

Macbride, 1910

%	f.t.		%	f.t.	
2.76	-32.8	(2+1)	48.38	74.7	(1+1)
7.86	-21.85		50.53	83.5	
13.14	0.02		52.37	86.5	
17.34	12.58		52.02	87.3	
19.78	18.78		53.41	90.4	
21.59	23.60		53.58	90.4	
22.65	27.23		56.45	97.0	
24.46	31.05		56.07	99.5	
29.29	40.90		57.01	99.5	
34.94	50.10		57.84	100.5	
40.36	60.03		60.09	104.1	
45.77	70.8		60.72	104.2	
46.44	70.15		58.97	104.7	
48.00	74.6		63.06	107.0	
48.38	75.2				
49.15	75.4		60.72	94.7	(2+3)
49.72	78.0		60.77	95.2	
50.37	78.7		61.93	106.4	
51.52	80.18		62.58	109.8	
52.40	82.5		63.06	113.6	
56.45	89.0		63.18	114.0	
57.01	90.8		63.57	115.7	
60.09	94.1		64.09	118.2	
			65.00	124.2	
			65.63	129.4	
			69.66	145.5	

Staronka, 1910

%	mol%	f.t.	
19.18	5.8	19	(2+1)
19.46	5.9	18.5	
30.45	10.2	39.5	
38.74	14.1	52	
51.20	21.4	74.5	
56.23	25	83	
60.69	28.6	98	
62.62	30.3	91.5	
63.60	31.2	92	
65.60	33.1	108	(1+1)
67.58	35.1	115.5	
70.70	38.5	130	
72.81	41	137	
74.56	43.2	142	
75.17	44	143.5	
77.71	47.5	159	(2+3)
81.17	52.8	173	

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-41.5	29.7	103.2
0.5	-42.8 E	30.4	105.2
1.1	-27.0	31.3	107.2
1.5	-23.0	31.9	112.2
2.3	-19.5	32.8	113.2
3.1	-17.0	34.6	126.9
3.8	-3.0	36.6	140.4
4.9	+9.4	40.3	148.3
8.9	32.6	41.5	150.3
11.1	43.5	42.5	158.2
13.1	52.0	43.6	162.6
16.3	62.8	46.5	171.6
19.6	74.1	50.1	182.5
22.4	81.0tr.t.	53.8	189.6
24.3	85.5metast.	59.2	194.8
25.9	90.7	62.6	195.5
26.1	90.8	71.7	190.8
27.4	93.3	73.3	188.1
29.9	96.9	76.1	183.7 E
24.3	87.5	77.1	192.7
25.9	97.5	79.0	201.0
27.4	99.6	82.0	214.5
28.3	101.7	100.0	277

(6+1),(2+1),(1+1),(2+3)

Herz and Martin, 1923

t	d	η
	15.077 %	
20	1.1102	1284.0
30	.0997	1090.0
40	.0892	946.0
50	.0788	828.5
60	.0683	740.7
70	.0572	661.3
80	.0470	599.8
90	.0364	542.0

Schönrock, 1900

%	d	(α)magn.
	16°	
17.5312	1.15228	0.9665
6.5675	1.03884	0.9705
0	0.98169	2.0085

Pearce, 1915, Anderson, 1915

N	λ		
	0°	25°	50°
2.0	0.009	0.036	0.045
1.0	0.019	.036	.030
0.5	.016	.021	.025
0.1	.016	.021	.027
0.01	.037	.061	.067
0.001	.130	.260	.400

Pyridine (C_5H_5N) + Mercuric bromide ($HgBr_2$)

Walden and Centnerzwer, 1906

%	D. b.t.	%	D. b.t.
3.76	+0.286	3.97	+0.308
8.03	.637	7.30	.594
11.42	.958	10.31	.873
14.38	1.285	13.13	1.180
17.36	1.647	15.73	1.468

Staronka, 1910

mol%	f.t.	E	
4.9	9	-	(2+1)
10	43.5	-	
12.5	57	-	
14.9	68	-	
19.7	69	-	
25.4	106	-	
25.4	115.5	-	
33.9	117	107	
39.5	108	106	
41.9	113	107	(1+1)
43.9	118	107	
46.2	121	108	
49.6	123	-	(2+3)
54.9	131	123	
58.8	134	-	
64.8	133	-	

Pearce and Anderson, 1915

N	λ 0°	λ 25°	λ 50°
2.0	0.012	0.034	0.043
1.0	.020	.026	.032
0.5	.018	.023	.026
0.1	.017	.023	.028
0.01	.031	.047	.053
0.001	.130	.280	.290

Pyridine (C_5H_5N) + Mercuric iodide (HgI_2)

Walden and Centnerszwer, 1906

%	D. b.t.	%	D. b.t.
		2.21	+0.140
4.26	+0.267	5.31	.341
7.87	.522	7.71	.509
11.58	.810	10.45	.712
15.07	1.104	13.13	.928
18.47	1.428		

Staronka, 1910

mol%	f.t.	E	
5	10	-	(2+1)
9.8	42.5	-	
15.14	66.5	-	
19.3	83	-	
26.3	102.5	-	
29.6	107	-	
34.6	107	-	
38	103	-	
43	97	86	
46.7	88.5	85	
48.5	89	87	1+1)
49	90	87	
49.8	90	87	
50.6	89	87	
51.3	93.5	87.5	
51.6	96	87.5	
52.7	108	-	
53.2	109	-	
55.4	122	-	
57.9	135	-	

Mathews and Ritter, 1917

%	f.t.		%	f.t.
1.93	-50	(2+1)	39.20	42.58
2.78	-42		41.00	45.10
4.27	-31.5		42.34	47.72
7.32	-20		43.15	50.02
10.29	-10		45.51	55.05
14.86	-0.1		46.96	57.07
18.42	+8.83		48.29	60.07
24.40	20.02		52.81	70.35
27.91	25.55		57.60	80.05
34.56	35.28		61.43	90.08
37.64	40.08		65.34	98.50

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.	
0.0	-41.5	35.0	113.7	
9.3	+36.8	49.3	88.5	E
11.4	49.1	53.4	111.7	
14.3	62.8	57.2	126.0	tr.t.
17.4	76.4	60.3	155.6	
22.4	95.3	100.0	250	
25.9	104.6			

(2+1)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	14.5-14.90°	
0	0.984	37.88
0.57	1.014	38.47
1.63	1.063	38.92
3.60	1.158	38.81
4.72	1.212	38.53

Sakhanov, 1913

N	λ
	25°
0.625	0.018
0.483	.018
0.279	.015
0.114	.014

Pearce and Anderson, 1915

N	λ		
	0°	25°	50°
2.0	-	0.013	0.018
1.0	0.009	.013	.018
0.5	.008	.012	.015
0.1	.013	.019	.024
0.01	.069	.102	.117
0.001	.266	.364	.448

Pyridine (C_5H_5N) + Mercuric cyanide (HgC_2N_2)

Walden and Centnerszwer, 1906

%	D. b.t.	%	D. b.t.
3.08	+0.350	2.46	+0.269
6.17	0.749	4.95	0.576
9.42	1.208	7.33	0.919
12.02	1.628	9.28	1.208
14.27	2.038	12.26	1.562

Staronka, 1910

mol%	f.t.	mol%	f.t.
7.1	9 (6+1)	22.3	42
8.7	11	22.9	45
10.1	12.3	23.7	46
10.4	12.2	25.3	53 (3+2)
11.3	13	26	54.5
12.9	13.5	26.6	56.5
13.8	14.5	27.5	68
15.8	16.5	27.7	70
15.9	20.5 (3+1)	29	86
17.3	22.5	32	111
18.4	28.5 (2+1)	33.8	122.5
19.3	32	34.4	125
20.6	38	38.3	141

Schönrock, 1900

%	d	(α)magn.
	16°	
0	0.98169	2.0085
23.2275	1.20198	0.4989
29.6018	1.28155	0.4685

Pyridine (C_5H_5N) + Ethylmercuric chloride ($Hg(C_2H_5)Cl$)

Walden and Centnerszwer, 1906

%	D. b.t.	%	D. b.t.
1.90	+0.201	2.06	+0.214
4.28	0.440	4.53	0.470
6.38	0.701	7.15	0.776
9.10	1.013	9.75	1.088
11.70	1.360	12.99	1.507

Pyridine (C_5H_5N) + Potassium mercuric iodide ($KHgI_3$)

Walden and Centnerszwer, 1906

%	D. b.t.	%	D. b.t.
1.99	+0.110	11.08	+0.605
5.09	0.270	14.48	0.852
7.83	0.427		

Pyridine (C_5H_5N) + Lead nitrate (PbN_2O_6)

Walton and Judd, 1911

%	f.t.	%	f.t.
2.84	-19.4	14.49	40.03
2.09	-14.5	18.05	45.00
1.86	-10.00	22.70	49.97
3.42	0.00	26.85	59.52
3.78	5.4	32.11 (3+1)	70.00
5.12 (4+1)	8.7	38.12	80.00
5.77	14.72	47.43	89.93
6.34	19.97	56.10	94.94
7.88	24.75	58.90 (2+3)	99.89
9.89	30.03	60.32	104.90
11.66	34.97	62.09	109.90

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-41.5	10.9	72.4
1.4	+25.0	12.8	76.2
2.0	40.0	19.7	93.7
2.4	42.3	22.4	98.9
3.3	50.9	24.6	102.0
5.7	62.8	25.3	103.6
6.8	65.1	25.7	132.6
8.8	69.3		

(4+1), (2+1)

Pyridine (C_6H_5N) + Cupric sulfate ($CuSO_4$)

Tombeck, 1900

t	p dissoc. (4+1)	t	p dissoc. (4+1)
0	2.5	80	385.0
20	7.5	90	547.5
58	142.5	100	770.0
70	260.0		

Pyridine (C_5H_5N) + Cobalt nitrate (CoN_2O_6)

Mitra and Sintra, 1950

t	p dissoc. (6-5)	(5-4)	(4-0)
23	5.7	-	-
27	6.9	-	-
30	8.0	5.0	3.7
35	10.2	7.4	5.6
40	13.0	10.3	8.1 (at 40.2)
45	16.2	14.6	11.7
50	20.2	19.8	16.0

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-41.5	8.6	81.2
0.3	+58.9	10.9	82.8
0.9	65.0	13.3	97.0
1.5	70.0	14.8	105.7
2.6	74.1	15.6	109.7
4.6	77.9	18.7	118.2
6.7	79.2		

(6+1)

Pyridine (C_5H_5N) + Aluminum bromide ($AlBr_3$)

Plotnikov and Balyasnii, 1931

%	f.t.	%	f.t.
96.2	85	81.4	59
90	83	80.2	93
87.5	50	79	98
86.8	45	70	88
85.7	- (4+1)	60	170
85.3	59		

1-Picoline (C_6H_7N) + Cuprous iodide (CuI)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-69.75	26.0	82.9
1.4	+25.0	28.6	87.0
2.4	40.0	32.4	92.0
5.1	58.8	36.3	104.7
11.3	68.7	38.4	116.5
14.6	73.1	39.5	145.0
17.9	76.2		

(2+1)

1-Picoline (C_6H_7N) + Silver iodide (AgI)

Tombeck, 1900

t	p dissoc. (3+2)	t	p dissoc. (3+2)
0	5.0	83	380.0
20	17.5	90	475.0
31	42.5	100	660.0
50	117.5	102	712.5
60	172.5	107	830.0
72	252.5	110	920.0

1-Picoline (C_6H_7N) + Silver nitrate ($AgNo_3$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-69.75	32.1	115.0
12.1	+25.0	38.7	122.0
13.4	33.0	44.8	125.0
16.3	49.7	47.5	125.0
20.2	72.0	50.7	138.0
25.5	99.5	-	159.8 tr.t.
27.2	106.9	100.0	208.6

1-Picoline (C_6H_7N) + Zinc iodide (ZnI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-69.75	16.6	118.7
1.3	+25.0	19.8	130.8
2.5	43.0	26.2	140.4
3.5	59.3	27.2	156.1
4.8	68.9	31.5	170.8
6.8	82.2	32.8	175.8
12.5	105.8		

(3+1), (2+1)

1- Picoline (C_6H_7N) + Zinc thiocyanate ($ZnC_2N_2S_2$)

Bhattacharya and Sinha, 1955

t	p dissoc.	t	p dissoc.
(2-1)			
50.0	4.2	74.0	10.0
55.0	5.1	81.5	13.0
61.0	6.4	87.2	15.9
65.0	7.3	91.0	17.6
67.5	8.0		

1-Picoline (C_6H_7N) + Cadmium iodide (CdI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-69.75	19.2	92.3
4.5	+25.0	21.3	98.9
6.4	40.0	24.4	108.5
9.7	57.2	28.6	116.5
15.2	78.5	33.5	189.5
17.6	86.5		

(2+1)

1- Picoline (C_6H_7N) + Cadmium thyocyanate ($CdC_2N_2S_2$)

Bhattacharya and Sinha, 1955

t	p dissoc.	t	p dissoc.
4-3			
40.0	5.9	65.0	20.0
45.0	7.6	71.8	27.5
49.7	9.9	76.5	33.1
54.5	12.4	85.2	48.4
60.0	15.9	90.0	58.8
3-2			
55.0	7.1	75.0	21.4
60.0	9.8	80.7	28.2
65.0	12.5	85.5	36.3
70.0	16.6	90.0	45.6
2-0			
65.0	6.0	85.0	11.0
70.0	7.1	90.0	12.6
75.0	8.1	95.0	14.5
80.0	9.8		

1-Picoline (C_6H_7N) + Mercuric chloride ($HgCl_2$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-69.75	14.7	66.0
4.6	+23.2	15.2	68.3
4.6	25.0	16.9	73.1
5.8	31.0	19.8	80.8
7.2	40.0	21.8	82.9
7.9	42.3	25.5	86.5
10.2	51.6	34.1	107.1
12.6	60.1	40.4	124.1
13.6	64.1	100.0	276.0

(2+1)

1-Picoline (C_6H_7N) + Mercuric iodide (HgI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-69.75	20.5	79.8
6.4	+29.0	24.1	101.6
8.6	40.0	30.1	117.5
10.6	47.0	33.6	126.0 tr.t.
12.3	51.7	37.8	132.5
13.2	55.2	46.5	147.0
16.1	61.9	100.0	250
17.9	63.9		

(2+1)

1-Picoline (C_6H_7N) + Cobalt thyocyanate ($CoC_2N_2S_2$)

Bhattacharya and Sinha, 1955

t	p dissoc.	t	p dissoc.
	4-3		
4.5	65.0	13.5	85.0
6.0	70.5	16.5	88.5
8.2	75.2	22.4	95.0
10.3	79.6		

1-Picoline (C_6H_7N) + Nickel thyocyanate ($NiC_2N_2S_2$)

Bhattacharya and Sinha, 1955

t	p dissoc.	t	p dissoc.
	4-3		
65.0	6.3	85.9	16.8
70.3	8.1	91.7	20.8
74.5	10.0	95.0	24.0
80.2	12.9		

Quinoline (C_9H_7N) + Mercuric bromide ($HgBr_2$)

Staronka, 1910

mol%	f.t.
4.4	88 (2+1)
8.9	111
14.3	127
17.6	134

Quinoline (C_9H_7N) + Mercuric iodide (HgI_2)

Staronka, 1910

mol%	f.t.	mol%	f.t.
4.7	100 (2+1)	37.7	160
9.1	115.5	41.6	165
13.2	133.5	43	165 (1+1)
23.1	138	46.1	167
26.7	145	48.8	170
27.2	145	49.5	169.5
29.8	151	54.4	166.5
31.4	153		
35.4	156		

Quinoline (C_9H_7N) + Mercuric cyanide (HgC_2N_2)

Staronka, 1910

mol%	f.t.	mol%	f.t.
4.2	45 (3+1)	13.2	137
6	54	17.4	161
8.2	61	22.5	180
	89	27.1	192
9.2	61		
	99	tr.t. 56°	

Morpholine (C_4H_9ON) + Sodium iodide (NaI)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-4.6	7.6	104.2
0.8	+25.0	7.6	104.7
1.3	40.0	6.7	107.2
2.2	46.7	5.4	108.7
2.8	50.3	3.9	120.0
3.4	53.8	3.7	122.5
3.5	66.0	3.5	138.4
3.7	84.0	3.4	141.4
3.9	91.0	2.8	144.4
5.4	98.7	2.2	155.2
6.7	100.2		

(6+1),(4+1)

Morpholine (C_4H_9ON) + Cuprous iodide (CuI)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-4.6	6.6	77.6
-	-4.9 E	9.6	82.3
0.5	+31.6	12.7	87.2
0.6	38.7	16.1	89.3
0.8	44.7	19.3	92.8
1.3	55.8	22.7	94.5
1.8	61.9	26.7	97.6
3.1	68.5	29.3	117.4
5.0	74.0	35.1	155

(2+1)

Morpholine (C_4H_9ON) + Zinc iodide (ZnI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-4.6	2.7	142.5
0.3	-5.2 E	4.4	167.1
1.1	+25.0	6.4	177.9
1.2	40.0	7.1	181.5
1.3	84.5	7.9	194.2

(6+1)

Morpholine (C_4H_9ON) + Cadmium iodide (CdI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-4.6	7.9	137.3
0.5	37.8	9.0	142.6
1.2	53.8	10.2	151.5
1.8	59.1	10.8	155.5
2.4	61.1	13.2	164.7
2.8	82.2	15.8	175.3
3.2	92.2	18.4	183.4
3.9	103.2	20.8	190.0
5.4	123.0	21.7	195.5
6.7	133.5		

(6+1) (4+1) (2+1)

Morpholine (C_4H_9ON) + Mercuric chloride ($HgCl_2$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-4.6	1.9	37.5
0.4	+44.7	3.3	94.0
0.8	60.9	5.0	98.0
1.3	78.2		

Triphenylarsine ($C_{18}H_{15}As$) + Triphenylbismuth ($C_{18}H_{15}Bi$)

Pascal, 1912

%	f.t.	m.t.	%	f.t.	m.t.
100	76	76	29.41	43.3	39.9
60	60	47	20	46	40
50	54.5	39.5	14.28	50.2	47.5
42.85	50.5	39.6	0	59	59
36.84	47.2	39.5			

Drew and Landquist, 1935

%	f.t.
0	60.5
25	44.3
100	78.3

Triphenylarsine ($C_{18}H_{15}As$) + Triphenylstibine ($C_{18}H_{15}Sb$)

Pascal, 1912

%	f.t.	m.t.	%	f.t.	m.t.
0	59	-	60	42	-
9.09	58	52.4	68.97	37.5	37.5
20	56.5	47	71.42	37.5	37.5
31.03	54.2	43	80	43	38.6
41.17	52	40	90.91	46.8	42
52.37	48.2	38	100	58	47.8

Drew and Landquist, 1935

%	f.t.
0	60.5
30	42.1
100	55

Triphenylphosphine ($C_{18}H_{15}P$) + Triphenylbismuth ($C_{18}H_{15}Bi$)

Pascal, 1912

%	f.t.	m.t.	%	f.t.	m.t.
0	79.1	79.1	62.50	48	41.6
25	66.5	56.2	71.43	54.3	49.6
40	56.2	45.5	80	62	58.5
45	50.5	42	90.91	70.5	68
50	46	42	100	76	76
55.56	42.5	42			

Drew and Landquist, 1935

%	f.t.
0	79.1
53	56
100	78.3

Tetraphenylsilicane ($C_{24}H_{20}Si$) + Tetraphenyltin ($C_{24}H_{20}Sn$)

Pascal, 1912

%	f.t.	E
0	233.0	-
9.09	231.3	229.8
25.92	228.0	225.2
37.50	225.7	223.1
50	223.3	222.0
39.68	222.1	221.2
67.27	221.2	221.0
75	222.1	221.0
83.32	223.3	221.1
90	224.6	222.0
100	225.7	225.7

Tetraphenylsilicane ($C_{24}H_{20}Si$) + Tetraphenyllead ($C_{24}H_{20}Pb$)

Pascal, 1912

%	f.t.	E
0	233.0	-
13.04	230.5	226.3
25.92	226.8	223.7
37.50	224.4	220.9
50	221.1	219.3
62.50	219.1	218.9
65.00	218.8	248.8
71.43	219.4	218.7
80	221.6	220.2
90.71	225.0	222.2
100	227.7	-

Tetraphenylsilicane ($C_{24}H_{20}Si$) + Tetraphenylgermanium ($C_{24}H_{20}Ge$)

Drew and Landquist, 1935

%	f.t.	m.t.
0.0	237.5	237.5
20.0	235.5	237.1
42.5	234.3	236.0
76.7	232.7	234.5
100.0	233.4	233.4

Formamide (CH_3ON) + Sodium iodide (NaI)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	2.55	18.4	31.7
2.3	-0.3	19.7	33.0
4.1	-2.1 E	21.7	40.0
4.5	+3.4	23.1	41.7
5.0	7.5	24.5	51.9
7.9	21.0	25.5	77.1
9.2	25.0	27.2	116.3
12.3	32.8	28.4	132.5
14.2	35.8	29.7	142.0
16.6	32.0	30.4	187.0
18.3	29.0 E	30.8	195.0

(6+1),(3+1),(2+1)

Formamide (CH_3ON) + Potassium iodide (KI)

Philippe and Lafontaine, 1958

%	f.t.
0.0	2.55
40.9	25.0
57.7	40.0

Formamide (CH_3ON) + Potassium thiocyanate (KSCN)

Philippe and Lafontaine, 1958

%	f.t.
0.0	2.55
49.2	25.0
53.5	40.0

Formamide (CH_3ON) + Thallium nitrate ($TlNO_3$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	2.55	34.7	113.2
3.4	7.0	37.2	113.2
4.6	19.3	37.6	117.2
6.8	39.2	42.5	123.0
10.0	56.8	45.4	125.7
14.7	75.6	47.2	127.7
18.2	86.5	48.3	128.2
21.5	94.9	50.1	131.5
26.2	101.3	50.3	132.5
27.2	102.5	54.6	137.5
28.9	104.7	63.7	144.6 tr.t.
29.8	107.7	64.8	148.0
30.2	109.7	77.4	173.3
33.5	112.2	100.0	206.0

(2+1),(3+2),(1+1)

Formamide (CH_3ON) + Cadmium iodide (CdI_2)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	2.55	14.7	84.9
0.2	2.5	16.2	86.5
0.4	1.7 E	17.2	86.5
0.7	5.4	18.4	93.9
1.3	14.8	20.1	98.5
2.1	28.4	22.5	106.4
2.9	36.3	24.6	110.2
4.0	45.4	26.7	105.2
5.1	52.8	27.5	102.0 E
6.1	59.4	28.7	109.2
7.7	65.0	31.1	135.4
10.3	76.8	36.0	169.6
		41.1	198

(4+1),(3+1)

Formamide (CH_3ON) + Manganese iodide (MnI_2)

Philippe and Lafontaine, 1958

mol %	f.t.	mol %	f.t.
0.0	2.55	17.3	109.2
1.6	- 5	21.0	113.7
3.5	-14	24.3	112
4.4	-18	28.5	106
5.9	-28	30.3	100.2 E
7.0	-35	37.1	108.2
7.8	-38 E	45.6	113.3
9.3	+47	51.4	115.3
10.0	62.9	57.3	123
10.8	72.5	62.3	147
11.6	86	71.1	183
12.9	88		

(4+1),(1+1)

Formamide (CH_3ON) + Thorium tetrachloride ($ThCl_4$)

Philippe and Lafontaine, 1958

mol%	f.t.
0.0	2.55
0.4	0.0 E
0.6	54.2
1.2	113.2
2.8	159.5
4.2	- Decomposition

Acetamide (C_2H_5ON) + Sodium bromide (NaBr)

Menshutkin, 1908

mol%(2+1)	f.t.	mol% (2+1)	f.t.
0	82	10.0	90
1.6	80	11.2	100
3.3	78	12.6	110
4.9	76	14.3	120
6.3	74	16.5	130
7.4	72	18.1	135
8.2	70	18.4	155
9.0	80	18.7	175

Acetamide (C_2H_5ON) + Sodium iodide (NaI)

Menshutkin, 1908

mol % (2+1)	f.t.	mol % (2+1)	f.t.
0	82	25.0	60
2.4	78	26.4	70
4.7	74	28.1	80
7.0	70	30.2	90
9.3	66	33.0	100
13.8	58	36.4	110
16.0	54	40.0	115
18.2	50	44.6	120
20.4	46	54.7	125
23.0	41.5	55.5	150
23.9	50	56.5	175

Acetamide (C_2H_5ON) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggeth, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
0	79-81	96.2	250
78.5	200	98.0	280

Acetamide (C_2H_5ON) + Potassium bromide (KBr)

Menshutkin, 1908

mol%	f.t.	mol%	f.t.
0	82	12.5	70
2.5	78	12.7	85
7.1	70	12.8	100
8.8	66	13.0	115
10.2	62	13.2	130
11.3	58	13.3	145
12.2	54	13.5	160
12.4	53 E	13.65	175

Acetamide (C_2H_5ON) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1908

% (6+1)	f.t.	% (6+1)	f.t.
3.1	30	62.5	100
12.4	75	65.0	110
21.7	70	68.0	120
31.0	65	71.5	130
40.0	60	75.5	140
48.5	55	80.0	150
56.0	50.5	82.4	155
56.0	50.5	85.5	160
56.8	60	90.0	165
57.8	70	92.7	167
59.0	80	100	169
60.5	90		

Acetamide (C_2H_5ON) + Magnesium iodide (MgI_2W)

Menshutkin, 1908

% (6+1)	f.t.	% (6+1)	f.t.
0	82	76.0	130
28.0	70	82.1	150
46.7	58	85.5	160
56.5	49	90.8	170
56.5	49	96.2	175
63.4	80	100	177
70.5	110		

Acetamide (C_2H_5ON) + Calcium chloride ($CaCl_2$)

Menshutkin, 1909

mol%	f.t.	mol%	f.t.
0	82	15.2	100
3.2	74	15.8	120
4.6	70	16.3	140
5.7	66	16.6	150
6.8	62	17.2	160
7.7	58	17.5	165
9.3	50	18.6	170
10.0	46	19.4	175
10.2	50	20.7	180
10.5	54	21.8	184
11.3	60	23.3	186
12.5	62	24.2	200
14.2	64	25.0	210
14.7	80	(6+1),(3+1),(1+1)	

Kusnezov, 1909

f.t. (4+1) 71 - 72°

Urea (CH_4ON_2) + Lithium nitrate ($LiNO_3$)

Howells, 1931

%	f.t.	%	f.t.
0	132.0	29.56	117.5
2.78	126.7	32.60	122.5
4.67	122.3	35.92	125.9
7.93	113.9	39.39	123.7
10.79	105.0	42.25	122.0
11.31	102.9	46.36	113.2
14.53	90.1	48.63	104.1
19.52	74.0	50.38	100.6
19.83	64.1	52.85	121.3
22.98	44.7	54.10	127.9
25.68	106.8	57.07	146.4
27.89	113.5		

(2+1)

Urea (CH_4ON_2) + Sodium nitrate ($NaNO_3$)

Howells, 1930

%	f.t.	%	f.t.
0	132.0	29.49	84.3 E
2.19	128.3	30.50	83.4
6.90	121.4	31.01	92.5
16.35	105.8	33.20	109.0
23.22	95.8	36.69	131.7
27.00	88.8	41.30	156.0
29.00	84.1		

Urea (CH_4ON_2) + Potassium nitrate (KNO_3)

Howells, 1931

%	f.t.	%	f.t.
0	132.2	26.25	114.1
2.52	128.9	27.79	120.1
5.50	125.2	29.61	124.9
8.93	121.2	31.09	129.8
14.75	115.6	32.04	136.1
21.79	110.9	34.26	141.8
24.60	110.1	36.50	151.6

Urea (CH_4ON_2) + Magnesium bromide ($MgC_4H_{16}N_8O_4Br_2$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	132	45.5	127
9.5	126	58.0	130
17.2	120	60.0	130
21.8	114	60.7	145
24.2	108.5	67.2	160
29.8	115	71.4	165
35.0	120	83.7	170
41.6	125	96.0	171

Urea (CH_4ON_2) + Calcium nitrate (CaN_2O_6)

Howells, 1931

%	f.t.	%	f.t.
0	132.2	24.00	102.4
3.90	127.0	26.88	117.7
5.25	125.1	27.58	119.9
7.45	121.6	34.55	143.1
12.62	111.7	40.13	151.3
18.04	100.5	44.02	147.2
20.04	90.7	45.60	144.6
23.18	96.3	47.20	152.6

(4+1)

Urethan ($C_3H_7O_2N$) + Magnesium bromide . Urethan ($MgC_{12}H_{28}N_4O_8Br_2$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	49	66.5	90
18.5	45	75.5	91.5
36.5	39	69.4	91
43.3	35	73.8	100
45.6	50	80.0	110
51.3	70	84.1	115
56.2	80	90.0	120
59.8	85	100	123

Urethan ($C_3H_7O_2N$) + Magnesium iodide . Urethan ($MgC_{18}H_{42}O_{12}N_6I_2$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	49	59.4	50
27.5	45	70.7	70
45.0	39	78.8	80
51.8	32	85.0	84
		100	87

Urethan ($C_3H_7O_2N$) + Mercuric chloride ($HgCl_2$)

Mascarelli, 1909

mol %	f.t.	mol %	f.t.
100	47.3	86.8	39.7
97.4	45.7	81.6	63.8
94.9	44.1	72.6	105.6
89.0	40.4	63.8	142.6
87.6	39.6		

Acetanilide (C_8H_9ON) + Magnesium bromide . Acetanilide ($MgC_{48}H_{54}O_6N_6Br_2$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	112	39.0	185
3.7	110	49.0	195
7.7	108	59.5	200
9.0	107.5	73.2	205
19.3	140	100	209
29.6	170		

Benzanilide ($C_{13}H_{11}ON$) + Antimony trichloride ($SbCl_3$)

Vanstone, 1914 and 1925 (fig.)

mol%	f.t.	mol%	f.t.
100	73	50	112 (1+1)
90	59	40	109
87.5	56 E	39	108
80	83	30	125
70	97	20	141
66.7	99	10	152
63	97	0	157
60	105		

Benzanilide ($C_{13}H_{11}ON$) + Antimony tribromide ($SbBr_3$)

Vanstone, 1914 and 1925 (fig.)

mol%	f.t.	mol%	f.t.
100	95.5	50	116 (1+1)
90	84	42	114 E
86.5	76 E	40	117
80	88 (1+2)	30	133
70	96.5	20	145
68	97 tr.t.	10	153
60	112	0	157

Methyl thiocyanate (C_2H_5NS) + Cobalt thiocyanate ($CoC_2N_2S_2$)

Gillis and De Sweemer, 1934

%	f.t.	E	%	f.t.
0	-53.55	-	0	22.17
6.6	-53.8	-	25	23.4
10.8	-54.5	-	35	25.2
12.8	-54.8	-	37	27.3
18	-57.2	-61.2	39	28.0
19	-58	-	41	28.8
21.5	-60	-	45	28.1
19.9	-59.2	-61.2	50	28.2
20.8	-59.8	-	55	28.2
			65	28.25
			70	28.3

Methyl thiocyanate (C_2H_5NS) + Aluminum bromide ($AlBr_3$)

Aléglane, 1940

mol %	κ
21.3	37.1
26.5	45.2
29.1	52.8
34.2	62.1

Ammonium thiocyanate (NH_4CNS) + Potassium thiocyanate (KCNS)

Wrzesnewsky, 1911-1912

mol%	f.t.	m.t.	tr.t. I	tr.t. II	tr.t. III
0	146	-	119.8	100	90.2
3	148.5	146.5	118.5	99.6	90.2
5	149.5	147.5	118	99	83.5
10	152	148	117	97.6	74
15	153.6	149.5	115.5	96.5	57
20	156	146.5	117	100	-
25	158	148	120.5	-	-
30	160	148	122	-	-
40	162	159.5	124.5	-	-
45	162.5	153	125.5	101	-
50	163	153.5	125	100	-
55	164	155	126.5	97	-
60	164	155.5	130	94.5	-
70	167	158	131.4	90	-
75	169.2	159.8	133	-	-
80	169.6	162	136	-	-
90	174	168	139	-	-
97	177	174.5	141.5	-	-
100	179	-	143	-	-

mol%	flowing pressure	mol%	flowing pressure
100	25	40	29.7
95	29.8	30	26.4
90	32.3	20	22.8
80	31.8	15	23.7
70	35.8	10	24.0
60	34.8	5	23.8
50	33.6	0	23.0

Nitromethane (CH_3O_2N) + Aluminum chloride ($AlCl_3$)

Galinker, 1956 (fig.)

%	d 20°	30°	40°
5	1.17	1.16	1.15
10	.18	.17	.16
15	.21	.19	.18
20	.24	.22	.21
25	.27	.25	.24
30	.30	.28	.27
35	.33	.31	.30
40	.35	.34	.32
45	.39	.37	.36

%	η (water=1) 20°	30°	40°
0	0.6	0.55	0.5
10	0.8	0.7	0.6
15	1.2	1.1	1.0
20	1.25	1.23	1.22
30	2.50	2.10	1.25
35	3.20	2.50	2.00
40	6.00	3.75	2.50
45	8.75	6.00	3.75

%	κ 20°	30°	40°
5	65	60	50
10	100	95	90
15	130	125	120
20	150	150	145
25	155	155	160
30	140	150	155
35	130	140	150
40	105	120	135
45	70	90	110

Nitromethane (CH_3O_2N) + Mercuric chloride ($HgCl_2$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	-28.6	8.9	159.0
0.3	-30.6 E	10.6	170.1
0.5	-13.0	12.9	180.4
0.8	+18.2	16.0	191.5
1.2	47.5	18.4	197.7
2.2	70.2	24.0	208.0
2.7	90.7	28.7	215.2
3.9	112.2	30.4	216.0
5.3	130.4	36.3	118.4
6.8	145.7	100	277

(1+1)

Nitrobenzene ($C_6H_5O_2N$) + Zinc salts

Martin and Pink, 1948

2^{nd} comp.		C.S.T.
Zinc laurate	$ZnC_{24}H_{46}O_4$	111.0°
Zinc myristate	$ZnC_{28}H_{54}O_4$	113.0°
Zinc stearate	$ZnC_{36}H_{70}O_4$	112.0°

Nitrobenzene ($C_6H_5O_2N$) + Mercuric iodide (HgI_2)

Mascarelli, 1906

%	f.t.	%	f.t.
0	5.67	5.20	114.6
0.324	5.90	6.27	121.3
1.10	50.51	7.40	129.3
1.68	69.63	8.35	133.9
2.20	79.1	9.34	141.8
3.35	95.3	10.71	149.3
4.44	105.8	13.27	163.3

Smits and Bokhorst, 1915

mol%	f.t. I	f.t. II	spontan. crystal.
0.73	97.9	-	-
1.02	108.0	-	106
1.32	118.1	116.8	113
1.56	124.8	123.9	120
1.63	126.5	126.2	121.7
1.74	128.6	128.9	122
1.81	130.7	131.1	127
1.87	-	133.1	127
2.46	-	144.3	139
3.07	-	155.0	149
3.90	-	165.8	163.5
4.63	-	173.7	-
5.42	-	180.4	-
8.03	-	197.0	186
12.04	-	214.6	210
18.81	-	229.2	226
23.43	-	235.9	231
49.57	-	244.7	242
72.62	-	245.1	242
86.81	-	247.0	-
94.41	-	250.4	248.6
100	-	255.3	-

Nitrobenzene (C_6H_5ON) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E	%	f.t.	E
0	6	-	68.6	25	-15
22.0	1	-15	72.5	35	"
37.4	-4	-15	76.6	45	"
48.0	-9	-15	80.9	55	"
55.3	-15	-	85.3	65	"
57.3	-17	-	89.9	75	-
58.3	-5	-15	94.7	85	-
61.5	+5	"	97.5	90	-
64.9	15	"	100	94	-

Nitrobenzene ($C_6H_5O_2N$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E
0	6	-
11.7	2	-16.5
20.4	-2	"
26.8	-6	"
32.0	-10	"
35.9	-14	"
39.2	-18	-
38.0	-16.5	-
40.5	-13.5	-16.5 (1+1)
44.0	-10.5	"
50.0	-7.5	"
56.0	-6.5	"
64.8	-6	-
67.5	-6.5	-
65.0	-15	-
67.2	-5	-6.5
69.6	+5	"
72.2	15	"
75.2	25	"
78.7	35	"
82.8	45	"
87.4	55	"
92.8	65	"
96.6	70	-
100.0	73	-

Nitrobenzene ($C_6H_5O_2N$) + Antimony triiodide (SbI_3)

Pushin, 1948

mol%	f.t.	E	mol%	f.t.	E
100	170	-	50	158	5
89	165	-	30	156	8.5
80	161	-	20	151	8.7
70	158	5	10	137	8.5
60	158	5	0 (1+1)	9	-

Nitrobenzene ($C_6H_5O_2N$) + Bismuth tribromide ($BiBr_3$)

Pushin, 1948

mol%	f.t.	tr.t.	E
100	218	148	-
89.5	202	133	-
80	190	125	-
70	172	116	-
62.5	162	98	-5
50.5	144	-	-4
40	127	-	-
35	113	-	-
30	100	-	-2
25	90	-	-3
15	61	-	-5
0	9	-	-

Nitrobenzene (C_6H_5ON) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.		%	f.t.
0	5.5		53.9	88
4.5	4		55.6	82
10.3	2		58.0	72
14.5	10	(2+1)	59.9	62
18.0	15		61.6	52
22.5	20		62.7	70
30.5	25.5		64.0	90
32.1	35	(1+1)	65.6	110
34.2	45		67.7	130
36.6	55		70.4	150
39.5	65		72.4	160
43.1	75		75.2	170
48.0	85		80.1	180
49.8	88		89.8	190
52.0	90		100	194

Plotnikov and Podornan, 1934

%	ϰ	%	ϰ
	18°		
0.08	0.31	11.4	14.27
0.45	0.99	13.6	14.61
2.31	4.24	17.3	13.33
2.35	4.26	17.7	13.73
4.9	8.35	21.8	11.36
6.4	10.47	22.9	10.57
7.3	11.37	26.4	8.06
7.8	12.86		

Vladimirova and Fotiev, 1955

%	ϰ	%	ϰ
	25°		
5.41	15.5	11.80	20.9
7.55	19.1	14.30	20.8
10.05	20.1	18.30	19.1

Nitrobenzene ($C_6H_5O_2N$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0.0	5.5	71.3	80
18.0	0	72.6	70
28.8	-5	73.9	60
36.3	-10	75.2	50
42.0	-15	76.4	40
44.3	0 (1+1)	77.6	30
45.7	10	78.9	20
47.4	20	80.7	30
49.4	30	82.4	40
51.6	40	84.1	50
54.0	50	85.8	60
56.7	60	87.6	70
60.0	70	89.8	80
63.6	80	91.6	85
66.1	85	94.3	90
68.4	87	96.6	93
70	85	100.0	96

Fialkov and Chor, 1949

mol%	D f.t.
0.0587	-0.032
2.13	1.175
3.39	2.176
7.35	4.304

Plotnikov, 1910

%	κ	%	κ
	18°		
1.9	3.65	19.0	21.4
2.4	5.05	20.8	21.2
2.95	5.80	20.4	21.3
3.05	5.80	23.1	20.1
3.56	6.60	26.3	18.6
3.85	7.05	28.6	16.90
4.4	8.30	28.6	17.00
4.95	8.75	33.2	11.90
5.55	9.95	36.6	10.20
6.1	10.90	39.9	6.89
8.95	15.2	42.0	5.41
10.7	17.3	46.7	2.74
11.4	17.6	49.3	1.47
14.4	20.0	53.3	0.57
16.7	20.9	59.9	0.078

Plotnikov and Bendetski, 1927

%	d	%	d
	18°		
1.14	1.217	28.3	1.477
1.60	.224	41.1	.621
9.3	.263	52.7	.748
16.1	.344		

%	κ 10°	κ 18°	κ 25°	κ 50°
0.27	-	-	0.71	-
0.37	1.14	1.10	1.07	1.09
0.61	-	1.43	-	-
0.82	-	1.86	-	-
1.14	2.42	2.52	2.11	1.96
1.60	-	4.00	3.79	-
2.0	-	-	4.38	-
2.2	-	5.34	4.80	-
3.0	-	-	5.83	-
3.3	6.27	6.38	6.29	5.23
5.0	-	9.80	-	-
5.8	10.23	10.90	10.68	9.61
6.6	-	12.08	12.24	-
7.2	-	12.98	13.31	-
9.3	14.93	15.59	15.90	12.26
13.2	-	19.50	20.1	-
16.1	19.26	21.14	22.3	20.94
17.2	19.19	21.71	22.9	22.26
21.1	-	20.9	-	-
21.3	-	-	21.78	24.34
24.4	-	18.40	20.80	-
28.3	13.9	16.64	19.45	25.20
30.4	-	-	18.55	-
32.1	-	-	16.70	-
32.4	9.51	13.11	16.39	23.22
41.1	-	6.21	9.06	-
42.5	2.70	5.03	7.68	16.0
44.8	1.56	3.28	5.64	14.6
47.3	0.86	2.21	4.10	13.0
51.4	0.24	0.89	2.02	-
52.7	-	-	1.60	-
58.6	0.014	0.107	0.35	3.02

Jacober and Kraus, 1949

mol%	λ
	25°
0.338	6.94
1.143	3.57
6.416	3.14
16.33	3.28
38.11	3.36

Nitrobenzene ($C_6H_5O_2N$) + Ammonium Aluminum bromide ($Al_3Br_{10}NH_4$)

Bigich and Sakhanovskaya, 1953 (fig.)

%	d 50°	40°	30°	20°
5	1.19	1.20	1.23	1.25
10	.24	.25	.26	.28
15	.27	.28	.29	.32
20	.32	.33	.34	.36
30	.40	.42	.43	.45
40	.51	.52	.53	.54
50	.63	.65	.67	.69
60	.80	.82	.83	.86

%	η 20°	30°	40°	50°
5	2000	1500	1000	500
10	2500	2000	1500	1000
20	4000	3000	2000	1500
30	6000	4000	2500	2000
40	13000	7000	4000	3000
45	21000	11000	6000	5000
50	42000	17000	9000	6000
53	76000	28000	14000	9000
59	-	6000	24000	13000
60	-	-	27000	14000

%	κ 20°	30°	40°	50°
5	18.0	20.0	23.0	25.0
15	40.0	45.0	51.0	58.0
20	45.0	52.0	60.0	67.0
25	45.0	54.0	63.0	72.0
31	41.0	52.0	64.0	75.0
41	31.0	42.0	56.0	70.0
50	18.0	28.0	39.0	51.0
55	10.0	18.0	28.0	40.0
60	5.0	10.0	19.0	28.0

N.B. For d, the order of temperatures of the curves in the original paper is erroneous

Nitrobenzene ($C_6H_5O_2N$) + Sodium chloride.Aluminum bromide ($AlBr_3.NaCl$)

Bigich, 1946

%	d 10°	20°	30°	40°
5.46	1.2633	1.2528	1.1704	1.1608
10.12	.2905	.2804	.2703	.2615
15.68	.3575	.3271	.3164	.3066
20.15	.3751	.3678	.3562	.3459
24.62	.4186	.4066	.3968	.3870
29.17	.4659	.4546	.4438	.4329
32.63	.5003	.4884	.4770	.4666
37.71	.5586	.5464	.5344	.5240
43.78	.6322	.6195	.6037	.5966
48.25	.6900	.6775	.6642	.6525

%	η 10°	20°	30°	40°
5.46	2977.5	2346.7	1871.7	1529.4
10.12	3668.7	2829.3	2213.8	1799.0
15.68	4658.5	3448.8	2638.7	2116.2
20.15	6136.0	4409.2	3317.8	2558.3
24.62	7414.0	5391.0	3959.6	3007.3
29.17	10707.0	7019.7	4967.9	3650.5
32.63	13960.5	8859.1	6108.9	4401.9
37.71	21759.3	12776.3	8320.8	5790.7
43.78	42743.4	22519.1	13387.1	8979.5
48.25	69014.3	33657.1	19215.9	12149.4

%	κ 10°	20°	30°	40°
5.46	22.4	26.4	31.0	36.1
10.12	35.3	43.0	51.2	59.4
15.68	42.7	51.8	62.9	73.9
20.15	43.1	54.0	68.9	76.8
24.62	44.0	55.0	68.5	82.5
29.17	39.1	53.1	68.2	85.1
32.63	35.8	49.9	65.3	83.9
37.71	27.4	40.9	56.3	72.7
43.78	18.0	29.8	44.2	59.7
48.25	12.7	22.3	-	-

Nitrobenzene ($C_6H_5O_2N$) + Potassium-Aluminum bromide ($KAlBr_4$)

Bigich, 1956,(fig.)

%	d 20°	30°	40°	50°
3.99	1.2242	1.2154	1.2054	1.1921
9.97	1.2579	1.2520	1.2450	1.2298
16.47	1.3375	1.3263	1.3128	1.3011
20.20	1.3685	1.3498	1.3296	1.3221
23.34	1.3876	1.3744	1.3661	1.3581
27.70	1.4300	1.4186	1.4097	1.4012
31.63	1.4775	1.4659	1.4542	1.4435
36.88	1.5500	1.5300	1.5100	1.4950
43.36	1.6204	1.6011	1.5790	1.5661
45.66	1.6528	1.6476	1.6243	1.6075
49.04	1.6903	1.6708	1.6600	1.6476
54.69	1.7860	1.7607	1.7457	1.7366
60.70	1.8883	1.8686	1.8523	1.8406

%	η 20°	30°	40°	50°
3.99	2000	1900	1800	1700
10	2100	2000	1900	1800
20	3000	2000	1900	1800
27.7	5100	3000	2000	1900
37	10000	5000	3000	2000
45	21000	11000	6000	4000
50	32000	15000	7000	5000
55	54000	26000	13000	8000

%	ϰ 20°	30°	40°	50°
5.08	6	7	8	9
10	12.5	14	16	17.5
20	17.5	21	23	27
22	17.5	21.5	24	28
27	17	21	25	29
31	16	20	24.5	29
40	12	14	22	26
50	6	.9	13	18
58.50	2.5	5.5	9	13

Nitrobenzene ($C_6H_5O_2N$) + Aluminum phosphorus chloride ($AlPCl_8$)

Fialkov and Buryanov, 1953 (fig.)

N	ϰ 25°	45°
0.13	20	26
0.25	33	44
0.50	49	65
0.77	64	87

Nitrobenzene ($C_6H_5O_2N$) + Antimony Aluminum bromide ($SbAlBr_6$)

Gorenbein, 1941

%	d 10°	20°	30°	40°	50°
63.9	-	2.0992	2.079	2.060	2.041
58.3	1.989	1.970	1.952	1.935	1.917
52.6	-	.855	.838	.822	.807
45.9	1.750	.735	.720	.705	.690
43.7	.715	.699	.686	.671	.656
34.6	.576	.565	.552	.538	.525
29.7	.513	.500	.488	.474	.462
26.7	.475	.463	.451	.439	.428
18.3	.381	.370	.359	.348	.337
14.5	.342	.331	.320	.310	.300
9.9	.298	.288	.277	.267	.257
7.9	.282	.272	.262	.252	.242
6.7	.268	.259	.249	.239	.229
3.6	.243	.233	.223	.213	.203
3.01	.237	.228	.218	.209	.199
0	.213	.204	.194	.185	.175

%	η 10°	20°	30°	40°	50°
63.9	-	-	37801	17903	9783
58.3	104855	41490	19804	10604	6418
52.6	-	21115	16503	6815	4369
45.9	20646	11540	7023	4510	3095
43.7	16999	9833	6159	3841	2822
34.6	8549	5576	3831	2716	2023
29.7	6482	4438	3154	2311	1778
26.7	5601	3902	2843	2119	1642
18.3	4141	3179	2496	2077	1691
14.5	3681	2851	2309	1869	1584
9.9	3195	2525	2068	1705	1457
7.9	3015	2403	1973.1	1651	1387
6.7	2908	2329	1925	1599	1376
3.6	2694	2183	1824	1525	1317
3.01	2633	2133	1771	1506	1295
0	2433.5	2006.3	1686.5	1420.3	1235.6

%	ϰ 10°	20°	30°	40°	50°
63.9	-	3.96	6.92	10.44	13.44
58.3	3.87	6.92	10.52	14.16	17.10
52.6	-	10.22	13.97	17.20	19.68
45.9	9.89	13.61	17.04	19.60	21.08
43.7	10.67	14.34	17.56	19.73	20.83
34.6	13.76	16.65	18.52	19.44	19.31
29.7	14.31	16.57	17.76	17.94	17,24
26.7	14.33	16.19	17.06	16.97	16.15
18.3	12.44	13.17	13.12	12.47	11.40
14.5	10.78	11.13	10.87	9.58	11.01
7.9	6.94	6.92	6.59	6.11	5.57
6.7	5.71	5.66	5.39	4.99	4.57
3.6	3.35	3.34	3.14	2.95	2.76
3.01	2.87	2.85	2.69	2.50	2.31

Nitrobenzene ($C_6H_5O_2N$) + Ferric chloride ($FeCl_3$)

Philippe and Lafontaine, 1958

mol%	f.t.	mol%	f.t.
0.0	5.70	24.3	35
1.9	3.8	35.2	38
3.3	2.5	37.0	36
5.8	0.0 E	42.0	32
6.5	2.5	42.8	28.5
9.2	10.0	44.6	27.0 E
11.6	16	45.0	35.5
13.3	18	46.2	53
15.5	24	46.6	98
19.6	29	51.6	126
		58.5	157
(2+1)			

Nitrobenzene ($C_6H_5O_2N$) + Ferric phosphorus chloride ($FeCl_8P$)

Fialkov and Buryanov, 1953 (fig.)

N	κ 25°	κ 45°	κ 65°
0.10	21	25	32
.25	35	44	55
.50	54	72	92
.57	60	80	102

Nitrobenzene ($C_6H_5O_2N$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al., 1942

mol %	f.t.	E
100	-23.5	-
85	+63.5	-
78	68.5	-
74.5	-	-25
61.5	74	-
54.5	75	-
51.5	75	-
48.5	75	-
41.5	73.5	2
31.5	69	5
22.5	62	4
12	40	5
0	9	-
	(1+1)	

Nitrobenzene ($C_6H_5O_2N$) + Stannic chloride ($SnCl_4$)

de Carli, 1929

%	d 15°	η 15°	η 25°
100	2.2398	1193	1101
95	2.1548	1285	1146
85	2.0129	1597	1332
70	1.8146	2065	1687
60	1.6979	2283	1827
50	1.5946	2427	1948
30	1.4142	2432	1986
10	1.2685	2342	1933
0	1.2046	2159	1799

Usanovich and Pichugina, 1956

mol%	d 20°	d 40°	d 60°	d 80°
100	2.2340	2.1819	2.1288	2.0790
82.80	1.9882	1.9408	1.8955	1.8596
73.27	.8535	.8178	.7763	.7389
70.66	.8257	.7857	.7485	.7108
59.38	.6953	.6604	.6213	.5899
41.07	.5315	.5016	.4707	.4435
40.58	.5101	.4801	.4539	.4272
21.34	.3517	.3273	.3050	.2796
0	.2080	.1864	.1680	.1485

mol%	η 20°	η 40°	η 60°	η 80°
100	928	758	637	540
70.03	1370	975	764	630
56.42	1590	1130	852	689
53.22	1680	1160	878	695
41.40	1920	1290	966	754
27.58	2110	1410	1040	806
24.39	2110	1420	1050	811
11.36	2100	1450	1090	840
0	2000	1410	1080	842

Nitrobenzene ($C_6H_5O_2N$) + Stannic bromide ($SnBr_4$)

de Carli, 1929

%	d 32°	η 32°	η 40°
100	3.319	2403	2154
80	2.418	2482	2156
60	1.936	2373	2022
50	1.760	2184	1934
40	1.609	2042	1800
30	1.484	1971	1703
20	1.373	1888	1602
0	1.192	1737	1481

Nitrobenzene ($C_6H_5O_2N$) + Hafnium tetrachloride. Phosphorus oxychloride ($HfOCl_7P$)

Larsen and Wittenberg, 1955

mol%	f.t.	mol%	f.t.
0	5.785	8.02	5.250
4.38	5.495	15.4	4.730

Nitrobenzene ($C_6H_5O_2N$) + Hafnium tetrachloride . Diphosphorus oxychloride($HfO_2Cl_{10}P_2$)

Larsen and Wittenberg, 1955

mol%	f.t.	mol%	f.t.
0	5.785	16.4	6.540
4.13	5.440	19.3	6.320
6.57	5.250	20.8	6.190
11.0	4.920		

Nitrobenzene ($C_6H_5O_2N$) + Zirconium tetrachloride. Phosphorus oxychloride ($ZrCl_7OP$)

Larsen and Wittenberg, 1955

mol%	f.t.	mol%	f.t.
0	5.785	6.74	5.325
3.77	5.555	9.13	5.155
3.92	5.525	9.44	5.170
4.52	5.495	11.0	5.025
4.59	5.490	11.0	5.035
5.77	5.405	11.2	5.015
6.05	5.385	23.5	4.205
6.14	5.375		

Nitrobenzene ($C_6H_5O_2N$) + Zirconium tetrachloride . Diphosphorus oxychloride ($ZrCl_{10}OP_2$)

Larsen and Wittenberg, 1955

mol%	f.t.	mol%	f.t.
0	5.785	15.0	4.580
4.58	5.385	22.6	4.060
6.71	5.220	23.5	3.960
11.8	4.840	43.4	2.560

m-Dinitrobenzene ($C_6H_4O_4N_2$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1910

%	f.t.	E	
0	90	-	
18.6	80	1	
31.3	70	1	
40.7	60	1	
48.0	50	1	
53.6	40	1	
58.0	30	1	
61.6	20	1	
64.5	10	1	
66.8	1	-	
68.8	-11	-	
52.5	27.5	-	
58.2	28.5	-	(1+1)
63.0	27.5	21	
67.5	25	21	
72.8	20	-	
76.2	15	-	
78.8	10	-	
80.8	5	-	
82.7	0	-	
64.9	-10	-	
66.8	1	-	
69.0	10	1	
71.6	20	1	
74.8	30	1	
78.7	40	1	
83.5	50	-	
89.0	60	-	
96.4	70	-	
100.0	73	-	

m-Dinitrobenzene ($C_6H_4O_4N_2$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1910

%	f.t.	E
0	90	-
15.7	85	47.5
29.1	80	"
40.6	75	"
50.0	70	"
57.3	65	"
63.0	60	"
67.4	55	"
70.8	50	"
72.0	47.5	-
73.4	50	47.5
75.6	55	"
78.2	60	"
81.0	65	"
84.0	70	"
87.2	75	"
90.4	80	"
93.6	85	"
96.8	90	"
100.0	94	-

o-Dinitrobenzene ($C_6H_4O_4N_2$) + Titanium tetrachloride ($TiCl_4$)
(author says m)

Hertel and Demmer, 1932 (fig.)

mol%	f.t.	mol%	f.t.
4.0	110	59	110.8
20	105	60	113
21.5	104.2	62.5	115 (2+3)
30	111	67	110.8
40	116.5 (3+2)	70	115.05
45	115	75	117 (1+3)
50.05	116 (1+1)	80	115

m-Dinitrobenzene ($C_6H_4O_4N_2$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	-23	-	55	60.6	-
95	+50	-	52	60.8	-
90	57	-24	50	60	-
85	59.5	-	43.7	64	57
80	60.5	-	33.3	70	51
75	60.2	-	23.5	75.5	48
70	60.5	-25	14	81.5	47
66.7	60	-	0	90	-
60	61	-			

m-Dinitrobenzene ($C_6H_4O_4N_2$) + Stannic chloride ($SnCl_4$)

Usanovich and Pichugina, 1956

mol%	f.t.	mol%	f.t.
92.58	29.5	37.90	66.5
88.15	39.0	28.55	71.5
87.53	39.5	25.65	73.5
74.99	50.5	22.71	76.0
66.39	55.5	19.50	76.5
51.51	60.0	0	87.5

%	d	η	d	η
	80°		100°	
100	2.0790	540	2.0320	471
82.29	1.9419	807	1.8940	706
62.21	.7678	1470	.7441	1090
47.82	.6673	1890	.6260	1440
31.28	.5647	2400	.5303	1820
9.02	.4350	2980	.4080	2170
0	.3928	3170	.3729	2310

o-Nitrotoluene ($C_7H_7O_2N$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	E	
0	-8.5	-	
11.3	-13.5	-	
18.5	-18.5	-	
21.3	-10	-18.5	
25.7	0	"	
31.1	10	"	
39.0	20	"	
44.0	25	"	
50.4	30	"	
55.7	33	"	
62.3	34.5	-	(1+1)
68.0	33	27.5	
72.3	30	"	
74.6	27.5	-	
70.0	10	-	
72.5	20	-	
75.5	30	27.5	
79.1	40	"	
84.5	50	"	
90.5	60	"	
97.5	70	-	
100.0	73	-	

o-Nitrotoluene ($C_7H_7O_2N$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	E	
0	-8.5	-	
10.5	-11	-13.5	
19.5	-13.5	-	
21.3	-10	-13.5	
27.6	0	"	
35.6	10	"	
41.0	15	"	
47.5	20	"	
55.7	25	"	
67.2	30	-	(1+1)
70.0	31	-	
72.4	32	-	
69.6	30	-	
73.5	40	-	
77.5	50	-	
81.7	60	-	
86.3	70	-	
91.4	80	-	
97.2	90	-	
100	94	-	

o-Nitrotoluene ($C_7H_7O_2N$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0.7	-8.5	49.3	99.5 (1+1)
1.0	-9.3	52.8	90
1.5	0	56.8	70
2.5	10	61.5	45
4.0	20	62.5	70
6.5	30	64.5	95
11.0	40	68.2	120
20.7	50	73.7	145
31.0	55	81.6	170
28.1	46	86.2	180
35.8	70	89.5	185
41.8	85	94.3	190
46.8	95	100.0	194

o-Nitrotoluene ($C_7H_7O_2N$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	-8.5	70.1	80
5.5	-10	72.0	70
8.7	-11	74.0	55
10.0	0 (2+1)	76.1	40
12.8	+10	79.1	19
17.6	20	82.5	40
24.8	30	84.9	55
33.0	40	87.5	70
47.7	42.5	89.7	80
54.3	60	91.3	85
59.5	75	93.8	90
63.1	85	95.8	93
66.0	90 (1+1)	100	96
68.8	85		

o-Nitrotoluene ($C_7H_7O_2N$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolie and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	-23	-	60	-60	-
87.5	-40.5	-	50	-62.5	-
79.5	-51	-	40	-60	-
71.5	-55	-	27.5	-47.5	-5
67.5	-	-26	0	-4	-

m-Nitrotoluene ($C_7H_7O_2N$) + Mercuric iodide (HgI_2)

Mascarelli, 1906

%	f.t.	%	f.t.
0	16.0	4.05	106.3
0.381	15.97	5.72	121.0
0.707	28.19	7.39	133.0
1.14	51.19	8.37	139.1
1.45	65.8	9.43	140.7
1.82	74.8	11.74	160.2
2.96	92.8		

m-Nitrotoluene ($C_7H_7O_2N$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	%	f.t.
0	16	76.3	30
15.0	10	80.8	40
30.7	0	86.0	50
39.2	-10	91.6	60
42.3	-20	98.0	70
67.2	0	100.0	73
72.5	20		

m-Nitrotoluene ($C_7H_7O_2N$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	%	f.t.
0	16	66.1	20
5.7	15	69.7	30
24.2	10	73.5	40
39.0	5	77.5	50
46.6	0	81.7	60
53.0	-5	86.3	70
56.8	-9	91.4	80
59.5	0	97.2	90
62.7	10	100	94

m-Nitrotoluene ($C_7H_7O_2N$) + Antimony triiodide (SbI_3)

Pushin, 1948

mol%	f.t.	E	mol%	f.t.	E
100	170	-	40	-	15
89.5	165	-	36	160	11
80	162	-	22	154	12
70	161	-	20	130	10
60	160.5	+7	0	15	-
50	160	10			

m-Nitrotoluene ($C_7H_7O_2N$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0.0	16	44.2	90
7.8	13 (2+1)	46.7	95
9.8	20	49.3	99.5
13.4	27	52.3	90
20.5	34	56.8	70
24.5	35	61.5	45
23.3	30 (1+1)	62.5	70
24.5	35	64.5	95
29.1	50	68.2	120
34.0	65	70.2	130
39.6	80		

m-Nitrotoluene ($C_7H_7O_2N$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	16	70.8	80
14.5	12	73.8	60
21.8	8	76.8	40
27.9	4	73.9	27
32.0	1	82.0	40
38.5	20	86.7	60
46.6	40	89.0	70
52.7	60	91.5	80
59.7	80	93.1	85
63.3	90	95.3	90
66.0	96 (1+1)	97.1	93
68.8	90	100.00	96

m-Nitrotoluene ($C_7H_7O_2N$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al; 1942

mol %	f.t.	E
100	-23	-
90	+57	-
87.7	58	-28
80	64	-
70	69	-
60	73	-
50	75	-
40	73	7.5
39	68	-
21.5	-	11.5
19	58	-
11.6	45	-
8.5	36	14.8
0	15.5	-
	(1+1)	

p-Nitrotoluene ($C_7H_7O_2N$) + Mercuric chloride ($HgCl_2$)

Mascarelli and Ascoli, 1907

%	f.t.	%	f.t.
0	51.34	8.22	104.1
0.56	51.19	11.19	128.4
1.12	52.9	13.78	141.1
2.89	74.3	15.73	152.4
7.21	101.3		

Mascarelli, 1909

mol%	f.t.	mol%	f.t.
0	51.3	7.5	141.1
0.28	51.2	18.6	192.5
0.57	52.9	43.2	241.5
1.5	74.3	52.5	248.3
3.8	101.3	40.0	277.6
4.3	104.1	100	280.0
6.0	128.4		

p-Nitrotoluene ($C_7H_7O_2N$) + Mercuric bromide ($HgBr_2$)

Mascarelli and Ascoli, 1907

%	f.t.	%	f.t.
0	51.54	3.75	58.47
0.78	51.36	4.36	61.77
1.01	51.30	4.99	63.77
1.44	51.24	6.17	74.77
2.18	51.09	11.35	112.2
2.46	51.05	15.99	133.0
2.82	52.42	25.70	158.5
3.13	53.22		

p-Nitrotoluene ($C_7H_7O_2N$) + Mercuric iodide (HgI_2)

Mascarelli, 1906

%	f.t.	%	f.t.
0	51.54	4.76	112.2
0.343	51.50	6.86	129.7
0.62	51.14	8.80	143.5
0.90	50.94	11.11	156.5
1.20	50.68	13.33	165.5
1.43	60.30	15.24	174.5
2.56	84.10	18.61	186.0
3.10	92.70	22.22	197.5
3.64	100.7	25.51	205.5
4.26	106.2	33.95	217.5

Smits and Bokhorst, 1915

mol%	f.t. I	f.t. II	spontaneous cryst.
0.83	99.3	92.3	-
1.06	107.8	103.8	-
1.47	119.6	118.3	-
2.08	132.9	134.1	-
2.29	136.8	138.7	-
3.60	-	159.3	-
5.67	-	179.2	-
9.92	-	205.6	203
16.21	-	227.3	222
18.97	-	233.2	-
23.69	-	238.2	234
26.31	-	242.5	-
33.00	-	247.0	244.5
41.49	-	249.6	248
82.96	-	250.8	248.5
89.57	-	251.2	249.5
100	-	255.3	-

p-Nitrotoluene ($C_7H_7O_2N$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1912

%	f.t.	%	f.t.
0	52.5	66.1	5
7.1	50	63.5	3
13.5	45	64.0	-20
26.9	40	67.9	0
33.6	35	70.0	10
38.8	30	72.5	20
46.0	20	75.5	30
50.6	10	80.0	40
54.4	0	85.0	50
57.3	-10	90.5	60
52.0	7.5 (1+1)	97.5	70
62.3	7.5	100.0	73

p-Nitrotoluene ($C_7H_7O_2N$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1912

%	f.t.	%	f.t.
0	52.5	68.2	20
11.3	50	71.6	30
29.3	45	75.1	40
42.2	40	78.9	50
50.0	35	82.9	60
56.3	30	87.2	70
61.0	25	92.0	80
64.7	20	97.5	90
67.0	16	100.0	94

p-Nitrotoluene ($C_7H_7O_2N$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0	52.5	53.4	100
9.2	47	58.3	80
15	42	61.7	60
19	37	64.0	45
23.6	50	66.5	75
29.1	65	69.5	105
34.8	80	73.5	135
41.3	95	80	165
46.2	105	86	180
49.3	109 (1+1)	94.3	190
51.7	105	100	194

p-Nitrotoluene ($C_7H_7O_2N$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	53.5	74.3	50
10.0	50	78.9	27
21.8	45	81.2	40
31.3	40	83.3	50
39.8	35	85.1	60
46.1	29	87.7	70
52.9	50	90.5	80
59.4	70	92.2	85
63.0	80	94.4	90
66.0	83 (1+1)	96.7	93
68.5	80	100.0	96
70.4	70		

p-Nitrotoluene ($C_7H_7O_2N$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	-23	-	48	72	-
97	+2	-	42	71	-
93	38	-	34.5	66	+40
90	46.5	-	28	58	45
85	55	-	26	56	45
82	59.5	-	19	45	45
75	64.5	-	16	-	45
70.5	67	-	13.5	-	-
68	-	-23	12	48.5	-
65.5	69	-	7.2	51	-
60	71	-	3.3	52.5	-
50	72.3	- (1+1)	0	55	-

m-Dinitrotoluene ($C_7H_6O_4N_2$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nicolic and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	-23	-	50	43	-
95	+31	-	45	45	-
91	40	-24.5	40	48	-
87	43.5	-	35	50.5	-
78	45	-	30	54.5	-
77	45	-24.5	27	56	-
70	46	-	18	61	-
65	46	- (1+2)	13	64	-
62	45	-	10	66	-
55	45	-	0	70	-

Trinitrotoluene s. ($C_7H_5O_6N_3$) + Antimony trichloride ($SbCl_3$)

Pushin, 1940-46

%	f.t.	%	f.t.
100	73	50	47
90	61	40	59
80	47.5	30	66
59	28	19.5	75
64.5	22.5 E	10	78
60	33	0	81

Trinitrotoluene s. ($C_7H_5O_6N_3$) + Antimony tribromide ($SbBr_3$)

Pushin, 1940-46

mol%	f.t.	E	mol%	f.t.	E
100	94	-	40	65	58
89.5	88	52	34	68	57
80	82	54	20	73	55
69	75	55	10	77	52
60	68	56	0	81	-
50	61	61			

Trinitrotoluene s. ($C_7H_5O_6N_3$) + Bismuth tribromide ($BiBr_3$)

Pushin, 1940-46

mol%	f.t.	E	mol%	f.t.	E
100	214	-	50	184	68
90	209	60	45	175	70
80	203	64	20	161	66
70	195	64	10	142	74
60	190	72	0	81	-

Trinitrotoluene s. ($C_7H_5O_6N_3$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al., 1942

%	sat.t.	f.t.	E
100	-	-23	-
92.9	-	+68.2	-
91.5	77	68.2	-
88.4	82	68.2	-
82.9	88	68	-
66.2	93.5	68.2	-
60	-	-	-23.5
56	92.5	68.2	-
45.5	-	-	-25.5
42.3	84	68.2	-
38	74	68.5	-
32.2	-	68.5	68
24	-	72.5	68.5
20	-	74.5	-
12.3	-	77	-
9.5	-	78	-
0	-	81	-

Trinitrotoluene s. ($C_7H_5O_6N_3$) + Stannic chloride ($SnCl_4$)

Pushin, 1940-46

mol%	f.t.	mol%	f.t.	E
100	-33	65.5	62.5	-33
97.5	+41	53	63	-
95	54	35	68	-
90.5	59	26	71.5	-
85	60.5	17	75	-
77.5	61.5	6.5	78.5	-
72	62	0	81	-

Trinitrotoluene s. ($C_7H_5O_6N_3$) + Stannic bromide ($SnBr_4$)

Pushin, 1940-46

mol%	sat.t.	f.t.	E
100	-		
95	-	71	-
90	81	70	28.5
85	93	71.5	28.5
79.5	98	72	28
70	101	71	28
60	102	71	28.5
50	101	72	28
43	96.2	71.5	28
37.5	93	73	27.5
31.5	83	72	29.5
28	77	71	27
22	-	73	21
11	-	77	-
0	-	81	-

1,3,4-Nitroxylene ($C_8H_9O_2N$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al., 1942

mol%	f.t.	E	mol%	f.t.
100	-23	-	50	64 (1+1)
90	36	-	45	63
80.5	46	-	40	62
70	-	-24	30	55
63.5	56	-	20	42
63	59	-	15.5	34
53	63	-	0	9

1-Nitronaphthalene ($C_{10}H_7O_2N$) + Mercuric iodide (HgI_2)

Mascarelli, 1906

%	f.t.	%	f.t.
0	58.00	4.05	105.85
0.32	57.90	5.11	116.35
0.77	57.80	6.22	127.85
1.43	59.95	8.09	143.35
2.04	74.35	10.27	153.35
3.07	91.35	11.70	161.35

1-Nitronaphthalene ($C_{10}H_7O_2N$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911-12

%	f.t.	E
0	57	-
13.6	50	-
21.0	45	-
27.3	40	-
32.0	35	-
35.8	30	-
39.0	25	- (1+1)
43.2	35	30
49.3	37.5	30
56.7	39	-
64.9	37.5	34.5
72.8	34.5	-
70.8	30	-
75.3	40	-
78.0	45	-
80.9	50	-
84.0	55	-
87.4	60	-
91.5	65	-
96.6	70	-
100.0	73	-

1-Nitronaphthalene ($C_{10}H_7O_2N$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911-12

%	f.t.	E	%	f.t.	E
0	57	-	55.9	35.5	33.5
7.3	56	23	62.6	37.5	-
23.2	50	"	67.6	38.2	-
34.8	45	"	68.0	38.0	-
42.6	40	"	68.8	40	38
49.2	35	"	73.4	50	38
54.6	30	"	78.4	60	"
61.5	23	-	83.8	70	"
64.7	30	-	89.6	80	-
45.3	31.5	-(1+1)	96.4	90	-
50.5	33.5	-	100	94	-

Nitrocellulose ($C_6H_{10}O_7N$) + Sodium nitrate ($NaNO_3$)

Urbanski, 1938

%	sound velocity (m/sec.)	%	sound velocity (m/sec.)
0	-	60	-
10	-	70	-
20	2080	80	-
30	-	90	-
40	2900	100	5400
50	3300		

o-Nitranisol ($C_7H_7O_3N$) + Stannic chloride ($SnCl_4$)

Pushin, 1948

mol %	f.t.	E
100	-33	-
93.5	0	-
82	+14	-34
76.5	+17.5	-
70	+20.5	-
60	+23	-
50	+23.5	-
42	+21.5	-
30	+14	-25
20	- 0.5	-28
18	-13.2	-
6	-10	-
0	+ 8	-
	(1+1)	

p-Nitroanisol ($C_7H_7O_3N$)+ Mercuric chloride ($HgCl_2$)

Mascarelli and Ascoli, 1907

%	f.t.	%	f.t.
0	51.81	9.35	90.0
2.28	51.11	12.16	110.1
4.76	59.7	15.25	125.1
6.84	79.0		

Mascarelli, 1909

mol%	f.t.	mol%	f.t.
0	51.8	5.51	90.0
1.30	51.1	7.26	110.1
2.75	59.7	9.23	125.1
3.99	79.0		

p-Nitraniline ($C_6H_6O_2N_2$) + Antimony tribromide ($SbBr_3$)

Kurnakov, Voskresenskaya and Gurovich , 1938

mol %	f.t.	mol %	f.t.
0	147.0	60	59.5
15	131.1	66.6	46.3
30	111.2	75	65.5
40	95.9	85	79.5
50	78.4	100	94

mol %	d	η
	95°	
40	2.1878	11922
50	2.4457	10523
60	2.6847	9136
66.6	2.8578	8343
75	-	7238
85	3.3221	5632
100	3.6912	3605

Guanidine nitrate ($CH_6O_3N_4$) + Lithium nitrate ($LiNO_3$)

Clark, Clow and al., 1949

mol%	f.t.	mol%	f.t.
0.0	214.0	44.5	155.2
16.7	196.0	48.8	166.0
24.8	185.8	56.1	186.0
31.9	175.0	80.0	236.5
36.7	166.5	87.5	244.6
44.3	155 E	100.0	259.5

Guanidine nitrate ($CH_6O_3N_4$) + Sodium nitrate ($NaNO_3$)

Clark, Clow and al., 1949

mol%	f.t.	mol%	f.t.
0.0	214	52.0	191
15.1	202	58.8	207
29.3	190	68.3	224
44.7	176 E	78.2	249

Guanidine nitrate ($CH_6O_3N_4$) + Calcium nitrate (CaN_2O_6)

Clark, Clow and al., 1949

mol%	f.t.	mol%	f.t.
0.0	213.8	32.8	135.0
4.0	206.3	38.1	120.8
8.1	201.0	38.2	120.5 E
17.3	177.3	40.8	134.2
23.6	159.1	42.7	143.7
29.1	143.4		

o-Chlornitrobenzene ($C_6H_4O_2NCl$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0	32.5	45.9	89 (1+1)
10.2	27	48.7	85
16.1	21	51.6	77
20.3	15	54.4	69
22.8	25	55.5	90
25.5	35	57.5	110
28.4	45	60.6	130
31.5	55	65.4	150
34.9	65	68.5	160
38.7	75	72.5	170
43.3	85	74.6	175

o-Chlornitrobenzene ($C_6H_4O_2NCl$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	32.5	67.0	70
11.2	30	69.5	60
21.3	25	71.7	50
29.4	20	73.7	40
35.9	15	75.7	30
37.5	13.8	77.5	21
39.5	20	78.9	30
43.1	30	80.6	40
46.7	40	82.3	50
50.3	50	84.0	60
53.9	60	85.8	70
57.6	70	88.6	80
61.5	80	92.4	90
62.9	83.5 (1+1)	95.8	93
64.2	80	100.0	96

m-Chlornitrobenzene ($C_6H_4O_2NCL$) + Aluminum) chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0	44.5	45.9	104 (1+1)
10.7	40	48.6	100
16.6	36	52.4	90
21.0	50	55.6	81
24.6	60	57.2	100
28.3	70	60.0	120
32.3	80	64.1	140
36.8	90	70.2	160
41.9	100		

m-Chlornitrobenzene ($C_6H_4O_2NCl$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	44.5	68.6	90
14.0	42	71.4	80
18.9	40	73.4	70
23.1	38	75.4	60
27.8	35.5	77.3	50
30.0	40	79.1	40
34.8	50	80.5	50
39.7	60	82.2	60
44.5	70	84.2	70
49.3	80	87.1	80
54.5	90	89.1	85
60.0	100	92.2	90
62.9	103.5	95.1	93
65.0	100 (1+1)	100	96

m-Chlornitrobenzene ($C_6H_4O_2NCl$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al., 1942

mol%	f.t.	E
100	-23	-
90	44	-
80.5	54	-
76	-	-23
70	58	-
60	60	-
50	61.5	- (1+1)
40	60	36
30	58	38.5
20.5	52	31
14	46	39
0	44	-

p-Chlornitrobenzene ($C_6H_4O_2NCl$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0	83.5	53.2	110
7.1	78	53.3	100
12.8	73	58.1	94
17.1	68	59.0	110
22.2	80	60.5	125
26.7	90	63.0	140
31.4	100	66.9	155
36.4	110	72.5	170
41.8	120	77.7	180
44.0	124	82.0	185
45.9	126 (1+1)	88.2	190
47.7	124	95.3	193
50.0	120	100.0	194

p-Chlornitrobenzene ($C_6H_4O_2NCl$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	83	69.6	80
9.0	80	72.4	60
24.8	70	75.2	40
36.6	60	78.0	20
41.0	70	81.7	40
45.6	80	85.3	60
50.2	90	87.2	70
54.9	100	89.3	80
60.1	110	93.2	90
62.9	115 (1+1)	95.4	93
64.8	110	100.0	96
66.8	100		

p-Chlornitrobenzene ($C_6H_4O_2NCl$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al., 1942

mol %	f.t.	E
100	-23	-
95	+19	-
94	26	-
91	33.5	-
83	42.5	-
76	48	-
71.5	-	-25.5
62.5	52.5	-
54.5	54	-
50	54.5	(1+1)
44	52.5	-
40	50	50
36	56	50
28	62.5	50
23.5	65	49
13	73.5	48.5
10	76	46
0	82.5	-

o-Bromnitrobenzene ($C_6H_4O_2NBr$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0	38.5	42.5	97
7.5	32	44.6	90
13.1	26	46.5	80
17.5	20	47.5	90
19.5	30	48.7	100
21.7	40	50.1	110
24.0	50	51.8	120
26.4	60	54.1	130
29.9	70	56.8	140
31.7	80	60.2	150
35.0	90	64.5	160
38.0	97	70.0	170
39.8	100 (1+1)	77.4	180

o-Bromnitrobenzene ($C_6H_4O_2NBr$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	38	68.6	40
19.7	30	70.9	30
30.0	21	72.0	24
33.8	30	73.4	30
37.6	40	75.5	40
41.5	50	77.6	50
45.3	60	79.8	60
49.2	70	82.4	70
53.0	80	86.3	80
56.9	88.5 (1+1)	88.6	85
59.7	80	91.9	90
61.9	70	94.5	93
64.1	60	100	96
66.4	50		

m-Bromnitrobenzene ($C_6H_4O_2NBr$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0	54.7	44.5	107
6.5	51	47.5	97
11.9	47	49.5	110
12.7	50	51.5	120
16.0	60	53.3	130
19.4	70	56.5	140
22.9	80	60.0	150
26.6	90	64.5	160
30.7	100	70.0	170
35.9	110	77.4	180
37.6	113	88.8	190
39.8	116 (1+1)	100	194
42.3	113		

m-Bromnitrobenzene ($C_6H_4O_2NBr$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	54	69.2	80
11.6	50	74.1	60
19.5	45.5	78.7	42
25.5	60	79.3	50
34.5	80	80.3	60
44.1	100	81.3	70
49.5	110	84.9	80
55.6	120	87.3	85
56.9	122 (1+1)	90.7	90
57.8	120	93.6	93
61.6	110	100	96
64.4	100		

m-Bromnitrobenzene ($C_6H_4O_2NBr$) + Titanium tetrachloride ($TiCl_4$)

Pushin, Nikolic and al., 1942

mol%	f.t.	E	mol%	f.t.	E
100	-23	-	49	70	-
90.5	+55	-	25	65.5	47
84	-	-23	14.5	60	48
80	64	-	12.5	51	-
70	69	-	5	53	-
60	71	- (1+1)	0	56	-
50	72	-			

p-Bromnitrobenzene ($C_6H_4O_2NBr$) + Aluminum chloride ($AlCl_3$)

Menshutkin, 1909-10

%	f.t.	%	f.t.
0.0	124.5	48.4	130
7.4	117	51.2	120
12.8	111	52.8	113
17.7	105	53.8	120
22.2	99	55.9	130
25.3	110	58.9	140
28.4	120	61.3	150
32.0	130	65.0	160
36.4	140	70.0	170
37.8	143	77.4	180
39.8	145 (1+1)	88.8	190
42.8	143	100.0	194
44.5	140		

p-Bromnitrobenzene ($C_6H_4O_2NBr$) + Aluminum bromide ($AlBr_3$)

Menshutkin, 1909

%	f.t.	%	f.t.
0	124.5	69.2	100
10.0	119	70.5	90
13.1	115	71.7	80
25.2	110	72.9	70
30.3	105	74.1	60
34.0	100	75.4	50
35.3	98	76.0	45
39.7	110	77.2	50
44.1	120	79.6	60
48.7	130	82.7	70
53.8	140	86.6	80
56.9	144 (1+1)	89.0	85
59.4	140	92.6	90
62.9	130	95.4	93
65.5	120	100	96
67.5	110		

Methyl borate ($C_3H_9O_3B$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d 25°	d 50°	η 25°	η 50°
0	0.9532	0.6231	361	281
10	1.0091	.7225	419	323
20	.2084	.9891	485	389
30	.4952	1.2096	573	452
40	.8254	.5678	695	566
50	2.0018	.8925	856	703
60	.4237	2.1954	1045	866
75	.9452	.7245	1442	1162
85	3.3012	3.1012	1835	1407
100	.3445	.2995	2572	1823

Ethyl borate ($C_6H_{15}O_3B$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d 25°	d 50°	η 25°	η 50°
0	0.7367	0.7039	557	414
10	.9343	.7564	623	482
20	1.0074	.9539	702	535
30	.2050	1.0030	791	625
40	.4154	.2034	902	729
50	.5138	.4022	1037	832
60	.7331	.6012	1183	987
75	2.2413	2.1618	1521	1221
90	.9345	.7534	2013	1531
100	3.3445	3.2995	2572	1823

LVI. METALLIC SALTS + HYDROXYL DERIVATIVES.

Methyl alcohol (CH_4O) + Lithium chloride (LiCl)

Lemoine, 1897

%	f.t.	%	d 21.5°	d 0°
26	1.0	5.2	0.854	0.836
27	23.0	14.5	.926	.910
30	50	22.1	.988	.974

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
31.1	0	30.5	30
30.7	10	30.6	40
30.6	15	30.7	50
30.5	20	30.9	60

Chéneveau, 1907

%	d	%	d
	14.5°		
17.49	0.9288	5.12	0.8348
13.11	.8933	0	.7976
9.41	.8655		

Schreiner, 1928

mol%	d	mol%	d
	18°		
0	-	3.117	0.90459
0.507	0.81336	4.062	.93572
1.005	.83170	5.136	.96882
1.993	.86656		

P.P. and N.S. Kosakewitsch, 1933

mol%	d	σ	mol%	d	σ
		20 - 21°			
0	0.791	22.40	4.08	0.830	24.03
2.15	.812	23.10	6.70	.856	25.12

Methyl borate ($C_3H_9O_3B$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d 25°	d 50°	η 25°	η 50°
0	0.9532	0.6231	361	281
10	1.0091	.7225	419	323
20	.2084	.9891	485	389
30	.4952	1.2096	573	452
40	.8254	.5678	695	566
50	2.0018	.8925	856	703
60	.4237	2.1954	1045	866
75	.9452	.7245	1442	1162
85	3.3012	3.1012	1835	1407
100	.3445	.2995	2572	1823

Ethyl borate ($C_6H_{15}O_3B$) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d 25°	d 50°	η 25°	η 50°
0	0.7367	0.7039	557	414
10	.9343	.7564	623	482
20	1.0074	.9539	702	535
30	.2050	1.0080	791	625
40	.4154	.2034	902	729
50	.5138	.4022	1037	832
60	.7331	.6012	1183	987
75	2.2413	2.1618	1521	1221
90	.9345	.7534	2013	1531
100	3.3445	3.2995	2572	1823

LVI. METALLIC SALTS + HYDROXYL DERIVATIVES.

Methyl alcohol (CH_4O) + Lithium chloride (LiCl)

Lemoine, 1897

%	f.t.	%	d 21.5°	d 0°
26	1.0	5.2	0.854	0.836
27	23.0	14.5	.926	.910
30	50	22.1	.988	.974

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
31.1	0	30.5	30
30.7	10	30.6	40
30.6	15	30.7	50
30.5	20	30.9	60

Chéneveau, 1907

%	d	%	d
	14.5°		
17.49	0.9288	5.12	0.8348
13.11	.8933	0	.7976
9.41	.8655		

Schreiner, 1928

mol%	d	mol%	d
	18°		
0	-	3.117	0.90459
0.507	0.81336	4.062	.93572
1.005	.83170	5.136	.96382
1.993	.86656		

P.P. and N.S. Kosakewitsch, 1933

mol%	d	σ	mol%	d	σ
		20 - 21°			
0	0.791	22.40	4.08	0.830	24.03
2.15	.812	23.10	6.70	.856	25.12

Cheneveau, 1907

%	n_D	%	n_D
14.5°			
0	1.3307	13.11	1.3626
5.12	.3434	17.49	.3741
9.41	.3536		

Schreiner, 1928

mol%	n C	n D	n F
18°			
0	1.32711	1.32871	1.33258
0.507	.33415	.33583	.33983
1.005	.34221	.34221	.34638
1.993	.35233	.35412	.35859
3.117	.36496	.36692	.37163
4.062	.37498	.37699	.38200
5.136	.38488	.38700	.39218

Methyl alcohol (CH_4O) + Lithium bromide (LiBr)

Kerler, 1894

%	D b.t.	%	D b.t.
1.56	+0.234	7.35	+1.29
2.97	0.272	10.20	1.885
4.57	0.725	11.89	2.34
5.09	0.842	13.82	2.81
7.39	1.271	14.28	2.9282

Gibson and Kincaid, 1937

%	d	%	d
25°			
0.000	0.78656	15.833	0.91760
4.033	.81742	24.352	1.00228
7.187	.84689	30.123	1.06676
11.711	.88100	35.114	1.12884
13.264	.89417		

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ	mol%	d	σ
12 - 13°					
0.00	0.797	23.30	6.75	0.928	26.41
1.94	.832	24.08	11.10	1.033	29.79
3.97	.876	25.10	14.23	1.078	31.50

Methyl alcohol (CH_4O) + Lithium iodide (LiI)

Kosakewitsch, 1928

mol%	d	σ	mol%	d	σ
14°					
0.00	0.798	23.30	5.31	0.917	25.83
0.53	.811	23.50	8.28	1.008	28.10
1.68	.835	24.01	12.68	1.105	30.81

Methyl alcohol (CH_4O) + Sodium bromide (NaBr)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
14.8	0	14.2	30
14.5	10	13.9	40
14.5	15	13.7	50
14.4	20	13.4	60

Gibson, 1937

%	d	%	d
25°			
0.000	0.78655	9.480	0.86139
3.097	.81022	13.368	.89473
6.257	.83504		

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ	mol%	d	σ
13°					
0.00	0.797	23.30	2.19	0.849	24.16
1.11	0.822	23.71	3.63	0.885	24.85

Gibson, 1937

%	π (1-1000 atm.)
25°	
0.000	-
3.097	76.72
6.257	73.82
9.480	70.99
13.368	67.66

Methyl alcohol (CH_4O) + Sodium iodide (NaI)

Lloyd, Brown and al., 1928

t	p dissoc. (3+1)
10	18.8
15	27.5
20	38.7

Woelfer, 1896

%	D b.t.	%	D b.t.
0.29	+0.032	6.30	+0.628
0.65	.071	6.79	.665
0.98	.101	7.87	.867
1.20	.125	8.28	.860
1.77	.182	8.65	.910
1.95	.194	10.37	1.120
2.10	.201	10.61	.160
2.17	.219	11.41	.275
2.58	.254	12.70	.481
3.34	.319	13.68	.623
3.50	.340	14.73	.828
3.78	.333	14.87	.833
4.46	.427	16.36	2.113
4.63	.450	16.50	.163
5.08	.494	18.84	.660
5.09	.490	21.13	3.26

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
39.4	10	44.9	30
42.2	20	44.7	40
43.8	25	44.6	50
44.3	27	44.3	60
45	28		

Gibson and Kincaid, 1937

%	d	%	d
	25°		
0.000	0.78656	22.774	0.92261
7.287	.84193	31.287	1.07766
14.520	.90325	38.209	1.16711

Kosakewitsch, 1928

mol%	d	mol%	d
	22°		
0.00	0.789	6.50	1.001
0.76	.815	7.08	1.002
1.48	.839	8.55	1.070
2.98	.888	12.53	1.194
4.33	.933		

mol%	σ	mol%	σ
	22°		
0.00	22.47	6.50	26.08
0.76	22.83	7.08	26.54
1.48	23.29	8.55	27.41
2.98	24.25	12.53	29.75
4.33	24.85		

Gibson, 1937

%	π (1-1000 atm.)
7.287	74.48
7.287	74.56
14.520	69.211
14.520	69.36
22.774	63.59
22.774	63.81
31.287	58.30
38.209	53.99
38.209	53.95

Methyl alcohol (CH_4O) + Sodium thiocyanate ($NaCNS$)

Hughes and Mead, 1929

%	f.t.	%	f.t.
25.9	15.8	33.8	48.0
28.6	24.7	34.0	48.9
31.1	34.6	34.8	52.3

Methyl alcohol (CH_4O) + Sodium methylate ($NaCH_3O$)

Le Blanc, 1889

%	d	n_D
	20°	
0	0.79223	1.32889
8.71	.85205	.34671
19.70	.92005	.36455

Thouvenot, 1913

%	d	(α)magn.
	25°	
0	0.7891	3.430
16.495	.8969	3.918
20.519	.9206	3.983
20.995	.9218	3.991

Methyl alcohol (CH_4O) + Sodium methyl cyanacetate ($NaC_4H_4O_2N$)

Thouvenot, 1910

%	d	(α)magn.
	25°	
0	0.7891	3.439
14.28	.8594	3.819
22.445	.8994	4.139
27.96	.9274	4.275
29.26	.9364	4.319

Methyl alcohol (CH_4O) + Sodium methyl camphor-carbonate ($NaC_{12}H_{17}O_3$)

Brühl and Schröder, 1904

%	t	d	n Hα	n D	n Hβ	n Hγ
0	14.50	0.7980	1.32948	1.33118	1.33490	1.33801
11.10	13.40	.8355	.34679	.34868	.35306	.35696
17.09	13.60	.8556	.35700	.35900	.36399	.36828
27.46	16.75	.8949	.37634	.37904	.38502	.39008
33.29	17.20	.9353	.39813	.40074	.40074	.41437

Methyl alcohol (CH_4O) + Potassium iodide (KI)

Timofeev , 1894

%	f.t.
12.7	11.4
12.8	12.2
13.8	13.5

Centnerszwer, 1910

%	f.t.	Crit.t.
8.64	-	266
12.95	0	-
14.2	20	-
14.6	-	262
14.97	25	-
19.2	85	256
26.8	115	242
28.9	144	229
29.6	188	196

Tyrer, 1910

%	f.t.	%	f.t.
12.66	15	23.49	180
13.94	30	22.54	200
15.90	50	21.57	220
18.36	80	19.87	240
20.00	100	18.43	245
21.38	120	17.35	247
22.60	140	12.12	250
23.43	160	7.06	252.5
		(retrogr. sol.)	

Solubility above the Critical point.

b.t.	% (V)					
c	10	20	30	36	40	45
252	0.3	0.99	3.57	7.06	10.55	15.32
270	-	-	3.38	6.89	10.31	-
280	-	-	3.29	6.80	10.15	-
290	-	-	3.29	6.71	9.91	-
300	-	-	3.19	6.54	-	-

Briscoe and Rinchart, 1942

molarity	d	η (alcohol=1)
	25°	
0.7523	0.8933	1.3695
.7899	1.3392	0.9013
.9812	.3586	.8779
1.2976	.3908	.8425
.4740	.4087	.8262
.7438	.4360	.8071
2.1770	.4804	.7876
	30°	
0.7482	0.8885	1.3608
.7879	1.3358	0.9065
.9784	.3549	.8839
1.2943	.3872	.8510
.4703	.4052	.8364
.7393	.4323	.8170
2.1716	.4767	.7978
	35°	
0.7442	0.8837	1.3544
	40°	
0.7403	0.8791	1.3496
.7840	1.3292	0.9176
.9737	.3484	.8985
1.2881	.3806	.8688
.4632	.3984	.8554
.7309	.4254	.8375
2.1620	.4696	.8224
	45°	
0.7365	0.8746	1.3461
	50°	
0.7322	0.8695	1.3433
.7800	1.3225	0.9293
.9690	.3418	.9119
1.2819	.3740	.8873
.4560	.3915	.8758
.7224	.4184	.8611
2.1516	.4626	.8464
	60°	
0.7762	1.3161	0.9418
.9641	.3350	.9293
1.2754	.3670	.9065
.4488	.3847	.8970
.7143	.4117	.8850
2.1412	.4555	.8725
	70°	
0.7722	1.3093	0.9550
.9592	.3283	.9445
1.2691	.3603	.9265
.4417	.3779	.9195
.7058	.4048	.9100
2.1309	.4486	.9000

Chatterji and Bose, 1950

%	t	η (alcohol=1)
9.09	35	1.261
	40	.250
	45	.248
	50	.251
16.67	40	1.459
	45	.453
	50	.448

Methyl alcohol (CH_4O) + Potassium thiocyanate (KCNS)

Stark and Gilbert, 1937

N	d
25°	
0.0000	0.786644
0.1811	.79964
0.3438	.80187
0.3918	.81419
0.8607	.84327
0.9545	.85141
1.4999	.88537

Chatterji and Bose,1950

%	t	η (alcohol=1)
17.16	35	1.964
	40	.942
	45	.910
	50	.907
23.78%	35	2.310
	40	.241
	45	.209
	50	.172

Methyl alcohol (CH_4O) + Potassium acetate ($KC_2H_3O_2$)

Kerler, 1894

%	D b.t.	%	D b.t.
2.45	+0.24	7.91	+0.938
3.50	.373	10.07	1.26
4.20	.415	10.14	1.273
5.14	.576	10.40	1.324
5.41	.585	11.47	1.473
6.14	.673	12.16	1.59
7.47	.888	13.39	1.806
7.54	.867	14.01	1.905

Methyl alcohol (CH_4O) + Potassium ethyl bortartrate ($KC_{16}H_{24}O_{12}B$)

Jones, 1933

%	d
20°	
32.38	0.9220
33.13	0.9238

λ	(α)	λ	(α)
20°			
32.38%		33.13%	
6708	-14.23	6708	-14.61
6363	16.26	6439	16.13
6104	17.89	6363	16.64
5893	19.67	6104	18.44
5782	20.50	5893	20.15
5780	20.57	5780	21.22
5700	21.43	5700	22.00
5461	23.97	5461	24.63
5218	27.13	5218	27.95
5153	28.05	5153	28.98
5105	28.84	4811	35.44
4811	34.51	4722	37.47
4722	36.53	4678	38.62
4358	47.09	4602	40.56
		4358	48.54

Methyl alcohol (CH_4O) + Magnesium chloride ($MgCl_2$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
13.4	0	15.1	40
13.6	10	16.0	50
13.8	20	17.0	60
14.3	30		

t	p. dissoc. (6+1)
10	34.4
15	48.6

Olmer, 1938

%	d	%	d
17°			
0	0.7895	15	.881
5	.8195	20	.911
10	.850		

Methyl alcohol (CH_4O) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

% (6+1)	f.t.	% (6+1)	f.t.
42.6	0	63.6	130
44.6	20	66.8	140
46.7	40	70.2	150
48.9	60	74.0	160
51.4	80	78.5	170
55.5	100	84.5	180
58.0	110	88.0	185
60.7	120	100	190

Methyl alcohol (CH_4O) + Magnesium iodide (MgI_2)

Menshutkin, 1907

% (6+1)	f.t.	% (6+1)	f.t.
49.6	0	66.2	120
52.6	20	69.5	140
55.3	40	73.2	160
58.0	60	77.1	180
60.6	80	81.5	200
63.3	100		

Methyl alcohol (CH_4O) + Magnesium nitrate (MgN_2O_6)

Lloyd, Brown and al., 1928

t	p.dissoc. (6+1)	t	p.dissoc. (6+1)
20	5.6	50	52.3
30	12.7	60	95.2
40	26.3		

%	f.t.	%	f.t.
13.6	10	18.9	40
14.8	20	21.2	50
17.3	30	25.9	60

Methyl alcohol (CH_4O) + Calcium chloride ($CaCl_2$)

Bonnell and Jones, 1926

t	p dissoc. (4-3)
15	49.30
20	69.88
30	99.3

Kerler, 1894

%	D b.t.	%	D b.t.
1.46	+ 0.098	7.02	+0.788
2.19	0.121	7.81	0.843
2.52	0.171	8.55	1.018
3.58	0.318	10.36	1.343
4.80	0.483	10.47	1.399
5.06	0.469	13.86	2.189
5.19	0.461	14.37	2.358
6.00	0.634		

Menshutkin, 1907

%(3+1)	f.t.	%(3+1)	f.t.	%	f.t.
33.3 (4+1)	0	60.5	55	55.7	190
37.6	10	63.1	75	57.7	215
42.2	20	66.3	95		
47.0	30	70.3	115	decomposition	
52.0	40	75.2	135		
57.3	50	81.8	155	above 230	
60.0	55	86.2	165		
61.3	56	89.5	170		
		93.5	174		
% in (3+1)		100	177		

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ	mol%	d	σ
		17.5°			
0.00	0.793	23.07	3.15	.881	24.78
0.43	.804	23.09	4.63	.916	25.76
1.23	.831	23.62			

Methyl alcohol (CH_4O) + Calcium bromide $CaBr_2$)

Bonnell and Jones, 1926

t	p dissoc. (4-3)
15	13.58
20	19.95
30	40.39
40	75.52

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
33.5	0	41.8	40
34.2	10	45.4	50
34.9	15	49.4	60
36.0	20	51.3	65
38.6	30		

Saxton and Lane, 1953

N	λ 0°	10°	20°	30°	40°
0.25	23.1	26.0	29.0	32.0	34.8
0.5	18.5	21.3	24.0	26.7	29.4
1.0	13.7	16.0	18.4	20.8	23.2
2.0	8.7	10.5	12.4	14.4	16.5
3.0	5.6	7.1	8.6	10.3	12.1

Methyl alcohol (CH_4O) + Calcium iodide (CaI_2)

Lloyd, Brown and al., 1928

t	p dissoc. (6+1)	(2+1)
20	47	20

%	f.t.	%	f.t.
53.8	0	56.7	30
54.9	10	57.7	40
55.3	15	58.7	50
55.9	20	59.7	60

Methyl alcohol (CH_4O) + Calcium nitrate (CaN_2O_6)

Woelfer,

%	D b.t.	%	D b.t.
0.90	+0.071	7.30	+0.488
1.78	.128	8.71	.592
2.17	.165	8.72	.588
2.99	.212	9.25	.631
3.38	.237	10.34	.711
3.82	.272	10.81	.741
4.65	.317	11.32	.775
5.12	.352	11.93	.832
5.85	.390	13.33	.957
7.27	.492	14.85	1.095

Lloyd, Brown and al., 1928

t	p dissoc. (2+1)
15	11.4
20	15.2
30	29.4
40	52.4
45	74.0

%	f.t.	%	f.t.
57.3	10	63.0	72
59.0	40	63.1	73
61.3	60	62.9	80
62.8	70		

Stark and Gilbert, 1937

molarity	d	molarity	d
	25°		
0.0000	0.786644	1.086	0.93528
.1705	.81113	1.175	.94724
.3215	.83230	1.583	.99874
.5344	.86163	2.025	1.05430
.5439	.88281	2.045	1.05755
.7487	.90063		

Saxton and Lane, 1953

N	λ				
	0°	10°	20°	30°	40°
0.25	9.85	11.0	12.3	13.7	15.5
0.5	7.20	8.05	9.22	10.3	11.5
1.0	4.88	5.65	6.42	7.30	8.20
2.0	2.40	2.95	3.55	4.25	4.85
3.0	1.37	1.75	2.20	2.70	3.25

Methyl alcohol (CH_4O) + Calcium lactate ($CaC_6H_{10}O_6$)

Henstock, 1934 (fig.)

%	f.t.	%	f.t.
11.5	15	52.5	55
20	25	57	65
31	35	b.t.	72
42	45		

Methyl alcohol (CH_4O) + Strontium bromide ($SrBr_2$)

Lloyd, Brown and al., 1928

t	p. dissoc. (3+2)	(2+1)
10	12.8	-
15	18.3	-
20	24.7	11
30	43.9	-

%	f.t.	%	f.t.
53.5	10	56.5	50
54.4	20	57.6	60
55.2	30	58.9	70
55.7	40		

Saxton and Lane, 1953

N	λ				
	0°	10°	20°	30°	40°
0.25	25.0	28.3	31.8	35.0	37.8
0.5	20.0	23.0	26.2	29.0	32.0
1.0	14.8	17.2	20.0	22.5	25.6
2.0	8.9	11.0	13.2	15.5	18.0
3.0	5.4	6.8	8.7	10.7	12.6
4.0	3.2	4.3	5.7	7.3	8.8

Methyl alcohol (CH_4O) + Barium bromide ($BaBr_2$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
30.8	0	29.0	30
30.1	10	28.7	40
29.9	15	28.3	50
29.5	20	28.0	60

Saxton and Lane, 1953

N	λ				
	0°	10°	20°	30°	40°
0.25	26.0	29.3	32.3	35.4	38.3
0.5	20.9	23.6	26.7	29.5	32.4
1.0	15.8	18.2	21.0	23.6	26.3
1.5	12.7	14.8	17.4	19.9	22.4

Humburg, 1893

%	d	(α)magn.
	16°	
0	0.7941	-
12.554	0.9125	18.494
17.847	0.9701	18.5939

Methyl alcohol (CH_4O) + Zinc chloride ($ZnCl_2$)

Dawson, Tockman and al., 1951 (fig.)

molarity	λ							
	-50°	-40°	-30°	-20°	-10°	0°	10°	20°
3.06	0.3	0.7	1.2	1.6	2.1	2.8	3.3	3.8
2.25	1.0	1.4	2.1	2.8	3.3	4.0	4.7	5.1
1.69	1.6	2.2	3.1	3.8	4.6	5.1	5.7	6.1
1.21	2.7	3.4	4.3	5.1	5.8	6.2	6.5	6.8
0.64	3.8	4.8	5.7	6.3	6.8	7.0	7.1	6.9
0.36	5.0	5.9	6.7	7.0	7.4	7.2	6.9	6.6
0.20	5.7	6.5	7.2	7.3	7.3	7.3	7.0	6.5
0.12	6.9	7.8	8.2	8.2	8.0	7.6	7.1	6.6

Methyl alcohol (CH_4O) + Zinc bromide ($ZnBr_2$)

Dawson, Toekman and al., 1951 (fig.)

molarity	λ							
	-50°	-40°	-30°	-20°	-10°	0°	10°	20°
3.39	0.2	0.3	0.8	1.2	1.5	2	2.3	2.8
2.79	0.4	0.8	1.2	1.7	2.2	2.8	3.2	3.6
2.25	0.8	1.3	2.0	2.5	3.0	3.5	4.0	4.3
1.80	1.5	2.1	2.8	3.3	4.0	4.5	4.8	5.0
1.21	2.4	3.1	3.9	4.4	4.9	5.2	5.4	5.5
0.64	3.8	4.5	5.1	5.5	5.9	6.0	5.9	5.7
0.36	4.8	5.4	5.9	6.1	6.1	5.9	5.8	5.3
0.20	5.7	6.3	6.5	6.5	6.4	6.2	5.8	5.3
0.12	7.2	7.7	8.8	7.7	7.3	6.8	6.2	5.8

Methyl alcohol (CH_4O) + Zinc iodide (ZnI_2)

Kosakewitsch, 1928

mol%	d	σ	mol%	d	σ
		22°			
0.00	0.789	22.47	2.79	0.991	24.23
0.90	.856	22.97	3.91	1.053	24.96
1.82	.922	23.61	5.07	1.127	25.84

Dawson, Tockman and al., 1951 (fig.)

molarity	λ							
	-50°	-40°	-30°	-20°	-10°	0°	10°	20°
3.06	0.3	0.8	1.0	1.3	1.8	2.1	2.4	2.6
2.10	0.8	1.2	1.7	2.2	2.5	2.8	3.1	3.3
1.80	1.3	1.9	2.4	2.9	3.3	3.6	3.8	3.9
1.21	2.4	3.0	3.5	3.9	4.2	4.4	4.5	4.5
0.56	3.5	4.9	5.2	5.4	5.5	5.3	5.1	4.8
0.30	6.1	6.5	6.7	6.7	6.6	6.3	5.9	5.6
0.20	7.4	7.8	7.8	7.7	7.4	7.0	6.5	6.1
0.12	10.9	11.2	11.0	10.7	10.1	9.5	8.7	8.1

Methyl alcohol (CH_4O) + Cadmium chloride ($CdCl_2$)

Lloyd, Brown and al., 1928

t	p.dissoc.		
	(3+1)	(2+1)	(3+2)
20	93.5	29	19

%	f.t.	%	f.t.
2.10	20	3.32	30
2.59	20	4.24	40

Methyl alcohol (CH_4O) + Cadmium bromide ($CdBr_2$)

Lloyd, Brown and al., 1928

t	p.dissoc.		
	(3+1)	(2+1)	(3+2)
10	52.2	-	-
20	94	25	19

%	f.t.	%	f.t.
8.98	0	17.4	30
11.9	15	19.7	40
13.9	20	23.7	50
16.5	25	30.5	60

Methyl alcohol (CH_4O) + Cadmium iodide (CdI_2)

Timofeev, 1894

%	f.t.
67.0	0
68.3	17
69.1	23

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
67.4	10	67.5	50
67.4	20	67.6	60
67.4	30	68.0	70
67.3	40		

Getman and Gibbons, 1915

N	%	d
	25°	
0.5	11.555	0.8766
0.2	4.602	.8269
0.1	2.301	.8099
0.05	1.150	.8021
0.02	0.460	.7969
0.01	-	-
0.005	0.115	0.7943
0.0005	0.012	0.7942
0.0	-	0.7942

P.P. and M.S. Kosakewitsch, 1933

mol%	d
17.8°-18°	
0	0.793
1.32	.863
4.35	1.023
9.33	1.279
12.71	1.444

Gibson, 1937

%	d	%	d
	25°		
0.000	0.78655	29.999	1.07068
3.331	.81056	41.559	.24029
5.767	.82908	50.628	.41309
11.661	.87752	59.332	.62681
20.993	.96657		

Sklyarenko and Smirnov, 1951-52

%	mol%	d 20°	d 30°	d 40°
10.06	0.9689	0.8708	0.8604	0.8503
25.45	2.900	1.0217	1.0099	0.9982
36.74	4.836	.1732	.1607	1.1479
49.57	7.920	.3808	.3674	.3532
54.92	9.600	.4892	.4747	.4597
63.37	13.00	.6974	.6818	.6662
65.49	14.24	.7783	.7612	.7456
68.65	16.07	.8833	.8664	.8507

Sklyarenko, Smirnov and Zhukov, 1952

%	mol%	d 30°	d 20°	d 10°
10.49	1.02	0.861	0.873	0.885
19.26	2.04	.942	.951	.966
26.72	3.09	1.017	1.028	1.038
33.00	4.13	.102	.110	.119
38.50	5.19	.178	.189	.196
43.33	6.27	.258	.268	.276
47.51	7.34	.334	.349	.358
51.51	8.50	.407	.422	.431
55.05	9.69	.484	.495	.507
58.25	10.91	.557	.570	.582
61.23	12.14	.630	.645	.655
63.92	13.42	.700	.719	.729
66.24	14.65	.779	.797	.807

%	mol%	d 0°	d -7°	d -15°
10.49	1.02	0.893	0.898	0.903
19.26	2.04	.972	.977	.983
26.72	3.09	1.049	1.054	1.058
33.00	4.13	.127	.135	.142
38.50	5.19	.208	.211	.222
43.33	6.27	.287	.292	.300
47.51	7.34	.368	.371	.381
51.51	8.50	.445	.454	.460
55.05	9.69	.523	.532	.542
56.67	10.29	.561	.572	-
58.25	10.91	.598	.611	.620
61.23	12.14	.676	.687	.700
63.92	13.42	.753	.763	.778
66.24	14.65	.831	.841	.856

Gibson, 1937

%	$Dv.10^4$ (1-1000 bars.)	%	$Dv.10^4$ (1-1000 bars.)
	25°		
5.767	77.94	41.559	61.44
11.661	75.70	50.628	56.38
20.993	71.70	59.332	51.43
29.999	67.44		

Sklyarenko and Smirnov, 1951-52

%	mol%	η (alcohol=1) 20°	30°	40°
10.06	0.9689	1.162	1.147	1.136
25.45	2.900	1.596	1.538	.491
36.74	4.836	2.184	2.091	.982
49.57	7.920	3.509	3.267	3.011
54.92	9.600	4.573	4.165	3.786
63.37	13.00	8.125	7.108	6.279
65.49	14.24	10.24	8.811	7.636
68.65	16.07	14.36	11.96	10.13

Sklyarenko, Smirnov and Zhukov, 1952

%	η(alcohol=1) 30°	20°	10°	0°	-7°	-15°
10.49	1.15	1.17	1.23	1.30	1.34	1.39
19.26	1.34	1.39	1.51	1.66	1.72	1.89
26.72	1.58	1.66	1.87	2.04	2.20	2.40
33.00	1.87	1.97	2.24	2.45	2.75	3.00
38.50	2.19	2.32	2.75	2.93	3.30	3.72
43.33	2.61	2.77	3.27	3.50	4.02	4.55
47.51	3.05	3.29	3.72	4.18	4.87	4.56
51.51	3.61	3.92	4.40	4.94	6.00	7.20
55.05	4.18	4.61	4.30	6.00	7.35	9.21
56.67	-	-	-	6.60	8.20	-
58.25	5.00	5.66	6.48	7.30	9.30	-
61.23	6.06	6.86	8.00	9.16	12.10	-
63.92	7.57	8.72	10.20	11.96	17.32	-
66.24	9.40	11.40	13.41	16.53	32.64	-

%	κ					
10.49	32.8	28.4	25.4	22.8	17.8	16.5
19.26	60.1	52.6	45.9	40.1	30.0	23.3
26.72	76.2	66.6	60.1	53.0	39.8	29.8
33.00	87.0	80.3	71.7	63.2	46.5	35.2
38.50	95.8	86.2	78.8	69.0	50.2	37.0
43.33	98.1	87.3	80.1	71.2	51.5	39.6
47.51	96.6	85.9	74.6	67.3	51.2	39.8
51.51	91.6	78.0	68.8	62.2	49.8	35.8
55.05	85.3	70.1	62.1	52.5	41.8	29.1
56.67	-	-	-	50.0	34.0	-
58.25	81.15	66.2	55.0	49.5	29.7	19.7
61.23	76.1	62.6	52.7	48.2	31.0	20.0
63.92	72.2	56.2	47.8	43.5	30.0	19.5
66.24	66.4	50.8	41.0	36.0	26.8	16.7

P.P. and M.S. Kosakewitsch, 1933

mol%	σ
17.8°-18°	
0	23.07
1.32	23.12
4.33	23.66
9.33	25.90
12.71	28.13

Getman and Gibbons, 1915

N	%	n_D
	20.5°	
2	-	-
1	-	-
0.5	11.555	1.34300
0.2	4.602	.33452
0.1	2.301	.33159
0.05	1.150	.33016
0.02	0.460	.32932
0.005	0.115	.32887
0.0005	0.012	.32887
0	-	.32875

N	%	λ
	25°	
2	-	4.14
1	-	5.40
0.5	11.555	6.20
0.2	4.602	6.82
0.1	2.301	7.14
0.05	1.150	7.49
0.02	0.460	7.89
0.01	-	8.39
0.005	0.115	9.33
0.002	-	11.58
0.001	-	15.14
0.0007	-	17.35
0.0003	0.012	18.82

Sklyarenko and Smirnov, 1951-52

%	mol%	ϰ 20°	ϰ 30°	ϰ 40°
10.06	0.9689	28.41	30.80	32.71
25.45	2.900	65.92	72.52	78.85
36.74	4.836	83.81	94.40	104.2
49.57	7.920	80.55	94.12	106.7
54.92	9.600	69.43	85.65	102.6
63.37	13.00	57.68	72.55	88.66
65.49	14.24	51.03	65.89	81.60
68.65	16.07	42.59	56.79	72.12

Saxton and Lane, 1953

N	λ 0°	λ 10°	λ 20°	λ 30°	λ 40°
0.25	5.70	6.28	6.70	7.10	7.38
0.5	5.23	5.73	6.20	6.67	7.00
1.0	4.40	4.90	5.38	5.88	6.23
2.0	3.14	3.60	4.05	4.50	4.90
3.0	2.12	2.52	2.95	3.33	3.71
4.0	2.75	2.39	2.04	1.70	1.37

Timofeev, 1905

mol%	U
20°	
0	0.600
0.99	.5720
9.1	.3563

Methyl alcohol (CH_4O) + Mercuric chloride ($HgCl_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (1+1)
15	44.7
20	59.2
25	87.6

Salvadori, 1896

%	D b.t.	%	D b.t.
0.99	+0.05	3.15	+0.13
1.10	.05	.29	.14
.23	.05	.37	.13
.28	.06	.96	.15
2.03	.09	4.89	.18
.24	.10	9.92	.35
.25	.10	15.42	.55
.57	.11	20.72	.755

Timofeev, 1891

mol%	f.t.
3.85	8.5
5.80	20
12.82	38.2

Timofeev, 1891-1894

%	f.t.	%	f.t.
10.6	-18	34.1	+19
11.1	-17	35.4	+20
25.4	+8.5	55.7	+36.5

Etard, 1894

%	f.t.	%	f.t.
7.6	-34	61.0	51
11.5	-20	63.6	62
12.8	-15	63.7	64
18.7	-2	64.3	74
23.2	+4	68.7	100
27.6	12	75.2	127
53.1	36		

Mc Intosh, 1896-1897

Below 38.8° (1+1) crystallises .

Stark and Gilbert, 1937

M	d	M	d
	25°		
0.0000	0.786644	0.3537	0.87070
.1092	.81264	.6931	.95145
.2713	.85120	.8493	.98862

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	21°	
0	0.791	23.01
0.94	.845	23.27
2.23	.921	23.66
3.89	1.015	24.40

Timofeev, 1891-1905

mol%	U
10-15°	
3.85	0.50776
1.96	.55406
0.99	.58236
0.50	.59703

Methyl alcohol (CH_4O) + Mercuric bromide ($HgBr_2$)

Timofeev, 1894

%	f.t.	%	f.t.
29.15	0	41.6	39
33.10	10	47.6	65
39.9	19	58.2	97
37.8	22		

Lloyd, Brown and al., 1928

t	p dissoc. (1+1)
10	52.0
20	74

%	f.t.	%	f.t.
34.8	10	41.9	30
39.5	20	43.2	40
41.3	22	45.0	50
41.4	27	46.0	60

Methyl alcohol (CH_4O) + Mercuric thiocyanide (HgC_2N_2)

Dukelski, 1907

%	f.t.	%	f.t.
26.10	0.0	32.53	31.7
29.17	14.7	33.29	38.1
32.01	23.4	34.05	44.5
31.77	27.4		

Methyl alcohol (CH_4O) + Cupric chloride ($CuCl_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (2+1)
10	36.0
15	50.7
20	68.0

%	f.t.	%	f.t.
36.1	0	38.2	40
36.5	10	39.2	50
37.0	20	39.9	60
37.5	30		

Etard, 1394

%	f.t.	%	f.t.
36.8	22	37.1	50
37.5	40	37.5	60

Saxton and Lane, 1953

N	λ				
	0°	10°	20°	30°	40°
0.5	4.88	5.42	5.90	6.37	6.72
1.0	3.32	3.70	4.12	4.54	4.94
2.0	1.99	2.31	2.70	3.10	3.50
3.0	1.32	1.61	1.92	2.28	2.63
4.0	0.93	1.19	1.47	1.79	2.13

Methyl alcohol (CH_4O) + Manganese chloride ($MnCl_2$), Nickel chloride ($NiCl_2$), Nickel bromide ($NiBr_2$)

Quartaroli, 1916

Specific magnetic susceptibility (see author)

Methyl alcohol (CH_4O) + Nickel bromide ($NiBr_2$)

Looyd, Brown and al., 1928

t	p. dissoc. (6+1)	t	p. dissoc. (6+1)
10	8.5	30	34.8
20	17.4	40	68.9

%	f.t.	%	f.t.
24.8	10	32.9	50
26.0	20	35.9	60
27.6	30	37.4	70
30.2	40		

Quartaroli, 1916

Specific magnetic susceptibility (see author .)

Methyl alcohol (CH_4O) + Cobalt chloride ($CoCl_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (3+1)	p dissoc. (2+1)	t	p dissoc. (3+1)
20	29.4	15	30	59.0
25	40.9	-	35	81.9

%	f.t.	%	f.t.
27.8	20	37.1	37
30.4	25	37.0	38
31.8	28	36.8	40
32.8	30	37.1	50
35.7	35		

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	17-17.5°	
0	0.793	23.07
0.89	.817	23.37
2.07	.862	23.83
4.40	.935	25.23
7.40	1.028	27.23

Quartaroli, 1916

Specific magnetic susceptibility.(see author.)

Methyl alcohol (CH_4O) + Cobalt bromide ($CoBr_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (6+1)	(3+1)	(2+1)
10	14.4	-	-
20	31.0	17	10
30	64.1	-	-

%	f.t.	%	f.t.
30.1	20	58.5	50
37.0	30	60.5	60
44.4	35	62.8	70
48.5	37	65.6	80
55.5	40		

Methyl alcohol (CH_4O) + Titanium chloride ($TiCl_4$)

Wertyporoch and Altman, 1934

M	κ	M	κ
	0°		
0.0460	32.0	0.8420	612.5
0.1270	197.0	1.14940	577.5
0.2260	220.0	1.18600	295.0
0.4410	382.0		

Methyl alcohol (CH_4O) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altman, 1934

M	κ	M	κ
	0°		
0.0868	62.0	0.6110	249.0
0.1293	69.5	0.7950	244.0
0.1715	79.8	1.1400	355.0
0.2550	96.0	1.4560	356.0
0.4160	166.4	1.7500	363.0

Methyl alcohol (CH_4O) + Ferric chloride ($FeCl_3$)

Quartaroli, 1916

Specific magnetic susceptibility (see author)

Methyl alcohol (CH_4O) + Molybdenum oxydichloride ($MoH_2O_3Cl_2$)

Vandenberghe, 1895

%	D b.t.	%	D b.t.
0.779	+0.070	13.94	+1.245
1.631	+0.151	16.41	+1.530
7.52	+0.615	16.77	+1.715
9.15	+0.850	17.95	+1.710
12.32	+1.060		

Methyl alcohol (CH_4O) + 4-Picoline zinc chloride ($ZnC_6H_7NCl_2$)

Flaschner, 1909

%	f.t.	sat.t.
89.0	86.4	-
78.5	73.5	-3.0
61.5	64.2	+1.5
49.7	60.5	+1.9
37.9	56.7	+1.1
25.2	52.0	-3.5
12.7	41.5	-
6.0	31.5	-

Ethyl alcohol (C_2H_6O) + Lithium chloride (LiCl)

Tower, 1908

g/l	p	g/l	p
		0°	
0	12.48	20.59	12.16
8.56	12.35	101.03	11.49

Skinner, 1892

%	b.t.	%	b.t.
0	78.43	8.01	82.61
2.4	79.13	9.93	83.98
5.39	80.58	15.94	90.18

Lemoine, 1897

%	f.t.	%	f.t.
14	1.6	14	25.0
14	5.7	15	40.6
13	13.0	18	62.6

Turner and Bissett, 1913

%	f.t.	%	f.t.	
12.60	0	19.60	20	
13.07	5	20.06	30	
14.36	10	20.24	40	(4+1)
15.82	15	19.61	50	
16.88	17	19.00	60	

Lemoine, 1897

%	d 0°	d 14.2°
0	0.809	0.797
5.2	0.851	0.839
10.1	0.881	0.871
14.6	-	0.903

Schreiner, 1928

N	d	n C	n D	n F
	18°			
0.0	-	1.35971	1.36143	1.36576
0.521	0.81208	.36689	.36874	.37325
1.008	.82955	.37272	.37464	.37928
1.537	.84767	.37860	.38056	.38534
2.015	.86324	.38351	.38555	.39045
3.043	.89462	.39277	.39494	.40005
0	-	1.35960	1.36132	1.36568
3.043	-	.39277	.39494	.40005

P.P. and M.S. Kosakewitsch, 1933

mol %	d	σ
	21°	
0	0.790	22.33
2.29	0.817	23.10
3.94	0.827	23.30
6.56	0.843	23.67
9.38	0.860	24.07

Gladstone and Hibbert, 1897

%	molar refraction H_α	molar refraction D	molar refraction H_β
	18°		
8.1	14.86	14.85	15.06
9.17	15.05	14.86	15.31
10.7	14.66	14.92	14.83
11.56	15.03	14.97	15.24
11.76	14.71	15.01	15.15
13.7	14.54	15.12	15.37
14.2	14.90	15.12	15.40
14.61	15.00	19.44	14.56

Ethyl alcohol (C_2H_6O) + Lithium bromide (LiBr)

Kerler, 1894

%	D b.t.	%	D b.t.
2.19	+0.231	7.87	+1.169
3.71	0.427	11.59	2.1
5.30	0.702	11.81	2.142
7.70	1.152	14.98	3.257

Bonnell and Jones, 1926

%	f.t.	%	f.t.
24.59	0	43.67	50
26.48	10	45.31	60
41.89	25	47.13	70
42.03	30	48.49	75
42.21	40	49.77	80

Ethyl alcohol (C_2H_6O) + Lithium iodide (LiI)

Kosakewitsch, 1928

mol%	d	σ	mol%	d	σ
		14°			
0.00	0.795	22.91	10.21	0.988	26.03
1.43	.821	23.11	14.22	1.057	27.44
2.40	.836	23.47	18.25	1.127	28.49
2.87	.847	23.56	19.47	1.147	28.87
2.96	.848	23.64	23.60	1.220	30.62
4.06	.872	24.23	26.92	1.278	31.95
5.08	.889	24.39			

Ethyl alcohol (C_2H_6O) + Lithium nitrate ($LiNO_3$)

Jones and Getman, 1904

N	D b.t.	N	D b.t.
0.06	+ 0.085	0.48	0.655
0.08	.11	0.51	.71
0.09	.120	0.58	.810
0.13	.17	0.71	1.010
0.18	.234	0.84	.21
0.19	.25	0.95	.415
0.27	.355	1.21	.87
0.35	.47	1.27	2.010
0.38	.515	1.46	.385
0.44	.44	1.67	.770

Campbell and Debus, 1956

%	d	%	d
	25°		
0.0	0.7850	12.43	0.8689
2.158	.7997	16.33	.8958
4.500	.8154	19.49	.9174
7.457	.8399	23.39	.9450

%	η (alcohol=1)	%	η (alcohol=1)
	25°		
0.0	1.000	12.43	3.215
2.158	1.276	16.33	4.365
4.500	1.591	19.49	5.617
7.457	2.209	23.39	7.51

%	κ	%	κ
	25°		
2.158	30.735	16.33	75.346
4.500	48.247	19.49	73.873
7.457	64.371	23.39	71.145
12.43	73.253		

Ethyl alcohol (C_2H_6O) + Sodium iodide (NaI)

Tyrer, 1910

Solubility above the critical point.

b.t.	% (V)						
c	10	20	30	36	40	45	48
262	0.5	2.06	4.85	7.92	10.39	14.75	17.76
270	-	-	4.85	7.71	10.31	14.53	-
280	-	-	4.76	7.58	9.91	14.16	-
290	-	-	4.67	7.41	9.50	-	-
300	-	-	4.58	7.15	9.25	-	-
310	-	-	4.49	6.98	8.84	-	-

%	f.t.	%	f.t.
10	30.45	27.80	220
30	30.67	26.58	230
50	30.79	24.64	240
80	31.03	20.76	250
100	31.08	17.35	255
120	31.12	9.75	260
160	31.03	7.92	261.5
180	30.70		
200	29.74		

Tammann and Hirschberg, 1894

%	d			
	0°	10°	20°	30°
0	1	0.98940	0.97881	0.96803
10.31	1	.99004	.98014	.97010
15.65	1	.99021	.98049	.97080
26.68	1	.99045	.98104	.97161

King and Partington, 1926

29.86% 25° d=1.037

Kosakewitsch, 1928

mol%	d	σ	mol%	d	σ
		24°			
0.00	0.786	22.22	4.54	0.892	23.52
.45	.804	22.42	6.02	.925	24.00
1.11	.816	22.72	7.96	.968	24.43
.80	.838	22.82	10.46	1.022	25.07
3.63	.882	23.41			

Haffner, 1903

c	d
15.56°	
0	0.7954
8.39	.8666
16.78	.9378
25.18	1.0090

c	η(alcohol=1) 15°	25°	35°
0	1	1	1
8.39	1.447	1.422	1.408
16.78	1.890	1.840	1.808
25.18	2.377	2.293	2.245

Ethyl alcohol (C_2H_6O) + Sodium thiocyanate (NaCNS)

Hughes and Mead, 1929

%	f.t.	%	f.t.
15.52	18.8	18.17	59.6
16.00	25.8	18.43	61.8
16.20	39.6	19.63	70.9
17.39	52.8		

Ethyl alcohol (C_2H_6O) + Sodium ethylate (NaC_2H_5O)

Le Blanc, 1889

%	d	n_D
	20°	
0	0.79501	1.36232
13.10	.85446	.37962
20.63	.88363	.38663

Ethyl alcohol (C_2H_6O) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggeth, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
1.3	30	87.5	200
2.3	50	94.8	250
9.7	90	97.0	280
69.6	150		

Ethyl alcohol (C_2H_6O) + Sodium ethylacetoacetate ($NaC_6H_9O_3$)

Brühl and Schröder, 1905

%	t	d
0	18.35	0.7917
8.800	18.05	.8219
13.54	15.40	.8407
15.70	17.00	.8468
22.56	18.30	.8704

%	t	n Hα	D	Hβ	Hγ
0	18.35	1.36062	1.36242	1.36662	1.37011
8.800	18.05	.37133	.37341	.37850	.38283
13.54	15.40	.37825	.38048	.38601	.39080
15.70	17.00	.38034	.38272	.38843	.39334
22.56	18.30	.38829	.39088	.39720	.40288

Thouvenot, 1910

%	d	$(\alpha)_{magn.}$
	25°	
0	0.78535	4.095
14.32	.8351	.492
16.43	.8406	.530
19.45	.8509	.658

Ethyl alcohol (C_2H_6O) + Sodium-1-lactate ($NaC_3H_5O_3$)

Purdie and Walker, 1895

g/100cc	N	$(\alpha)_D^{mol}$
23.208	-	-2.55
19.790	-	-2.49
11.200	1.0	-0.90
9.294	-	-0.54
7.467	0.66	+1.50
5.600	0.50	2.30
2.240	0.20	10.00
1.120	0.10	16.00
0.560	0.05	23.00

Ethyl alcohol (C_2H_6O) + Sodium Ethyl Camphorcarbonate ($NaC_{13}H_{19}O_3$)

Brühl and Schröder, 1904

%	t	$d_{25°}$
0	18.35	0.7917
12.74	18.40	.8293
14.20	17.40	.8349
23.00	17.60	.8618
28.71	15.25	.8825

%	t	n Hα	n D	n Hβ	n Hγ
0	18.35	1.36062	1.36242	1.36662	1.37011
12.74	18.40	.37678	.37923	.38445	.38885
14.20	17.40	.37960	.38193	.38713	.39167
23.00	17.60	.39186	.39422	.40027	.40551
28.71	15.25	.40123	.40383	.41036	.41610

Ethyl alcohol (C_2H_6O) + Potassium hydroxyde (KOH)

Guglielmo, 1882

%	t	d	%	t	d
0	17.0	0.79595	9.24	11.5	0.8755
1.408	13.4	.8113	12.61	14.0	.9008
3.732	13.2	.8303	14.42	10.7	.9192
5.535	16.3	.8431	18.36	12.6	.953
8.625	12.0	.8688	21.27	11.7	.9730

%	κ 0°	κ 12°	%	κ 0°	κ 12°
1.408	18.70	23.41	12.61	36.84	55.87
3.732	34.37	44.72	14.42	-	48.37
5.535	41.01	54.80	21.27	22.88	39.73
8.625	44.90	62.91			
9.24	40.43	59.77			

Ethyl alcohol (C_2H_6O) + Potassium acetate ($KC_2H_3O_2$)

Raoult, 1890

mol%	p 78°
0	760.0
1.912	745.3
6.695	711.2
10.209	687.9

Ethyl alcohol (C_2H_6O) + Potassium d-lactate ($KC_3H_5O_3$)

Purdie and Walker, 1895

g/100cc	N	$(\alpha)_D^{mol}$
30.700	-	-11.98
12.810	1.00	-13.80
5.124	0.40	-17.00
2.562	0.20	-19.50
1.281	0.10	-21.75

Ethyl alcohol (C_2H_6O) + Potassium oleate ($KC_{18}H_{33}O_2$)

Laing, 1918

%	D b.t.	%	D b.t.
4.32	+ 0.173	18.24	+0.820
9.26	.388	21.21	1.008
12.35	.528	22.59	1.100
13.57	.590	24.26	1.238
14.77	.654	26.90	1.350
16.78	.764		

N	d	κ
	60°	
0.68	0.8220	19.27
.49	.8030	18.12
.38	.7920	16.94
.31	.7850	15.66
.27	.7800	14.52
.22	.7750	12.85
.16	.7705	11.37
.11	.7640	8.91
.02	.7565	3.81

N	d	κ
	40°	
0.39	0.8013	14.77
.32	.7961	13.72
.27	.7922	12.91
.22	.7884	11.34
.17	.7848	10.31
.11	.7810	8.14
.026	.7755	3.42

Bhatnagar and Kapur, 1934

%	d	$(\alpha)^{magn.}$ (water=1)
	at room temper.	
19.35	0.8333	0.9348
18.57	.8309	.9249
16.40	.8283	.9241
15.01	.8247	.9200

Ethyl alcohol (C_2H_6O) + Magnesium chloride ($MgCl_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (6+1)
20	1.02
30	3.2
40	8.95

%	f.t.	%	f.t.
3.48	0.0	11.34	50
4.16	10	13.71	60
5.30	20	14.67	65
6.89	30	15.70	70
9.09	40		

Olmer, 1933

%	d
17°	
0	0.792
5	.817
10	.841
15	.868
20	.892

Ethyl alcohol (C_2H_6O) + Magnesium bromide ($MgC_{12}H_{36}O_6Br_2$)

Menshutkin, 1907

%	f.t.	%	f.t.
17.2	0	73.8	80
24.9	10	76.2	85
32.7	20	78.7	90
40.3	30	82.3	95
47.8	40	86.7	100
55.1	50	90.0	103
62.2	60	94.4	106
68.8	70	100.0	108.5
71.4	75		

Ethyl alcohol (C_2H_6O) + Magnesium iodide ($MgC_{12}H_{36}O_6I_2$)

Menshutkin, 1907

%	f.t.	%	f.t.
21.9	0	82.7	120
33.2	20	87.2	130
44.4	40	90.0	135
55.3	60	93.3	140
65.5	80	96.0	143
74.7	100	98.0	145
78.8	110	100	146.5

Ethyl alcohol (C_2H_6O) + Magnesium nitrate (MgN_2O_6)

Lloyd, Brown and al., 1928

t	p dissoc. (6+1)
10	2.4
20	6.2
25	9.7
30	14.5
35	22.5
40	32.3

%	f.t.	%	f.t.
1.449	0	14.19	50
2.979	20	19.50	60
5.114	30	25.38	70
9.796	40	24.60	80

Ethyl alcohol (C_2H_6O) + Calcium chloride ($CaCl_2$)

Raoult, 1890

mol%	p
78°	
0	760.0
1.966	743.2
5.156	710.9
7.393	676.2
10.788	639.3

Hayward and Perman, 1931

%	p	%	p
20°		30°	
2.710	43.72	0	78.65
5.222	-	3.207	77.78
6.987	42.69	6.262	77.09
9.532	41.34	8.210	76.08
11.68	41.16	10.75	74.61
13.32	40.19	12.87	72.99
13.54	40.13	15.31	71.24
15.95	38.94	16.39	69.94
18.08	37.43	18.79	67.43
19.91	36.03	20.50	65.37
		21.03	64.86
40°		50°	
0	133.7	0	221.6
3.514	132.7	3.013	220.2
4.704	131.5	6.392	216.6
6.542	-	8.995	212.7
10.46	126.6	10.85	210.4
16.35	118.4	13.53	205.4
18.80	114.7	16.32	200.3
20.56	111.5	18.37	194.0
25.20	103.0	20.90	187.6
26.67	99.56	23.99	177.6
		26.41	171.0
60°		60°	
0	351.5	18.80	309.8
3.249	348.9	19.67	305.0
7.628	340.6	22.68	292.1
10.18	334.4	24.71	278.2
15.54	319.3	28.32	258.0
15.77	-		

Bonnell and Jones, 1926

t	p. dissoc. (3+1)
30	7.74
40	17.81
50	48.98

Skinner, 1892

%	b.t.
0	78.43
6.705	79.13
11.5	79.88
16.85	81.13
24.3	83.63

Kerler, 1894

%	D b.t.	%	D b.t.
1.83	+0.153	2.03	+0.181
4.27	.453	4.85	.521
6.80	.763	7.49	.875
9.07	1.093	10.03	1.260
11.06	.419		
14.63	2.115		

Tammann and Hirschberg, 1894

%	d (0°=1)			
	0°	10°	20°	30°
0	1	0.98940	0.97881	0.96803
1.28	1	.98960	.97913	.96854
8.63	1	.99011	.98020	.97031
25.11	1	.99136	.98260	.97374

Menshutkin, 1907

% (3+1)	f.t.	% (3+1)	f.t.
34.8	0	86.8	30
46.0	20	89.2	85
58.7	40	91.9	90
73.0	60	96.2	95
80.8	70	100	97

Hayward and Perman, 1931

%	d	%	d
20°		30°	
2.710	0.81040	0	0.78036
5.222	.83014	3.207	.80566
6.987	.84188	6.262	.82761
9.532	.86270	8.210	.84416
11.68	.87877	10.75	.86430
13.32	.89230	12.87	.88080
13.54	.89400	15.31	.89921
15.95	.91395	16.39	.90924
18.08	.93576	18.79	.93026
19.91	.94730	20.50	.94428
		21.03	.94956
40°		50°	
0	0.77301	0	0.76300
3.514	.79945	3.013	.78587
4.704	.80768	6.392	.81250
6.542	.82600	8.995	.83254
10.46	.85736	10.85	.84738
16.35	.91120	13.58	.86923
18.80	.92840	16.32	.88938
20.56	.94037	18.37	.90833
25.20	.97390	20.90	.92948
26.67	.98940	23.99	.95590
		26.41	.96808
60°		60°	
0	0.75407	18.80	0.90811
3.249	.78099	19.67	.91153
7.628	.81169	22.68	.93850
10.18	.83447	24.71	.95880
15.54	.87485	28.32	.98947
15.77	.88126		

Kosakewitsch, 1928

mol%	d	σ	mol%	d	σ
		24°			
0.00	0.786	22.22	3.79	0.854	23.23
0.94	.808	22.62	7.47	.915	23.97
1.98	.824	22.88			

Chéneveau, 1907

%	d	n_D	%	d	n_D
		22°			
17.50	0.9264	1.3957	7.68	0.8474	1.3756
14.42	.8984	.3890	3.97	.8190	.3682
11.15	.8730	.3825	0	.7903	.3603
9.44	.8604	.3791			

Desai, Naik and Desai, 1934

N	λ	N	λ
30°			
3.2	1.71	0.01	22.78
1.0	4.00	0.001	50.78
0.1	7.50		

Vallender and Perman, 1931

%	Q diss.	%	Q diss.	%	Q diss.
20°		30°		40°	
17.76	-3.98	20.27	-5.46	22.66	-7.66
16.62	.51	19.52	.11	22.47	.26
15.82	.18	18.70	4.84	21.46	6.54
15.69	.02	17.86	.33	20.41	5.94
14.89	2.85	16.97	.04	19.43	.20
14.73	.71	16.12	3.42	18.17	4.63
13.96	.30	15.23	.03	17.22	4.63
13.12	.09	14.38	2.63	16.30	3.64
12.32	1.78	13.58	.42	15.46	.32
11.56	.51	12.74	.06	14.57	2.85
10.81	.43	12.01	1.90	13.75	.50
10.07	.18	11.28	.61	13.27	.40
9.35	.04	10.48	.39	11.96	1.87
8.65	0.86	9.90	.23	11.44	.70
8.00	.77	9.24	.06	10.52	.47
7.14	.61	8.40	0.87	10.40	.34
6.33	.51	7.48	.74	9.17	.11
5.64	.43	6.64	.60	8.45	0.88
4.78	.28	5.86	.50	7.23	.69
4.15	.24	5.16	.40	6.60	.56
50°		60°			
25.13	-9.88	26.69	-13.30		
24.30	.23	25.69	12.11		
23.37	8.64	25.44	11.45		
22.11	7.70	24.41	10.42		
21.13	.13	23.33	9.55		
20.86	6.86	22.14	8.05		
19.32	.20	21.21	7.44		
18.84	5.64	20.10	6.75		
17.95	4.77	19.03	5.87		
17.10	.27	18.01	4.93		
16.26	3.77	16.85	.49		
15.42	.43	15.78	3.63		
14.65	.02	14.85	.13		
13.89	2.62	13.97	2.87		
13.11	.25	13.13	.47		
12.28	.06	12.31	.05		
11.65	1.85	11.45	1.76		
10.88	.59	10.60	.57		
9.99	.25	9.54	.15		
8.67	.02	8.54	.02		
8.15	0.88				

Q diss. in cal/gr. alcohol .

Ethyl alcohol (C_2H_6O) + Calcium bromide ($CaBr_2$)

Bonnell and Jones, 1926

t	p dissoc. (3-1)
20	3.79
30	9.32
40	20.59
50	41.55
60	73.24

%	f.t.	%	f.t.
31.33	0	39.63	50
32.44	10	43.19	60
33.56	15	43.40	70
34.35	20	50.54	75
35.04	25	50.76	80
35.70	30	51.43	85
37.54	40		

Jahn, 1891

%	d	(α)magn. (current=1)
	20°	
20.37	0.99656	115.38
10.38	.88431	98.61
0	.79009	84.77

%	n Hα	n D	n D'	n Hβ
20.37	1.3980	1.4001	1.4002	1.4056
10.33	.3784	.3803	.3804	.3853
0	.3591	.3610	-	.3651

Ethyl alcohol (C_2H_6O) + Calcium nitrate (CaN_2O_6)

Raoult, 1890

%	p	%	p
60.20°		79.32°	
0	305.6	0.00	819.2
24.32	279.5	24.32	749.1

mol%	p
78°	
0	760.0
2.553	740.3
5.340	713.6
8.269	649.1

Lloyd, Brown and al., 1928

t	p dissoc. (2+1)
20	13.3
30	27.3
40	55.3

Jones and Getman, 1904

M	D b.t.	M	D b.t.
0.03	+0.04	0.37	+0.40
.06	.08	.45	.50
.08	.10	.47	.515
.14	.17	.57	.62
.19	.22	.66	.71
.21	.21	.74	.81
.26	.295	.83	.91
.30	.34		

M	D b.t.	M	D b.t.
2nd series .			
0.0428	+0.057	0.3380	+0.383
.0594	.078	.3486	.395
.1137	.135	.3815	.425
.1153	.138	.3978	.444
.1981	.239	.5361	.582
.2161	.253	.5630	.620
.2590	.300	.6862	.740
.2920	.333	.7029	.750
.2964	.340	.8274	.885
.3081	.351	.9118	.972

Lloyd and Jones, 1928

%	f.t.	%	f.t.
31.6	10	42.4	50
33.9	20	45.1	60
36.0	30	47.4	70
38.6	40	47.9	80

Saxton and Lane, 1953

N	λ				
	0°	10°	20°	30°	40°
0.5	0.67	0.83	0.99	1.17	1.34
1.0	.48	.62	.77	.93	1.10
1.5	.38	.50	.63	.79	0.96
2.0	.29	.39	.50	.64	0.80

Muchin and Farle, 1916

M	n_D	M	n_D
	18°		
0	1.3599	1.0	1.3823
0.2	.3651	1.5	.3926
0.5	.3720	2.0	.4108

Ethyl alcohol (C_2H_6O) + Strontium bromide ($SrBr_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (1+1)
10	3.0
20	15.2

%	f.t.	%	f.t.
38.9	10	42.9	50
39.0	20	43.0	60
39.4	30	43.1	70
42.4	40	43.8	80

Fonzes-Diacon, 1895

t	d
	39.2%
0-40	1.210 (5+2)

Jahn, 1891

%	d	(α)magn. (current=1)
	20°	
18.01	0.96351	105.86
10.10	.88132	95.13
0	.79009	84.77

%	n			
	Hα	D	D'	Hβ
	20°			
18.01	1.3872	1.3896	1.3895	1.3949
10.10	.3752	.3770	.3772	.3819
0	.3691	.3610	-	.3651

Saxton and Lane, 1953

N	λ 0°	10°	20°	30°	40°
0.5	2.86	3.40	3.91	4.40	4.38
1.0	2.43	2.91	3.50	4.10	4.63
1.5	2.10	2.57	3.21	3.80	4.40
2.0	1.76	2.23	3.31	3.46	4.09

Ethyl alcohol (C_2H_6O) + Barium iodide (BaI_2)

Bonnell and Jones, 1926

%	f.t.	%	f.t.
43.74	0	43.25	40
43.61	10	43.13	50
43.50	20	43.02	60
43.37	30	42.89	70

Ethyl alcohol (C_2H_6O) + Zinc chloride ($ZnCl_2$)

Chéneveau, 1907

%	d	n_D
	22°	
22.02	0.8964	1.3880
17.40	.8475	.3774
12.26	.8045	.3693
5.53	.7521	.3592
0	.7132	.3512

Blümcke, 1884

%	t	d	U
0	15	0.797	0.683
8.55	18	.856	.625
20.12	18	.949	.557
30.89	18.2	1.047	.493
40.40	17	.157	.448
44.20	18.5	.218	.428

Ethyl alcohol (C_2H_6O) + Zinc iodide (ZnI_2)

Kosakewitsch, 1928

mol%	d	σ	mol%	d	σ
		24°			
0.00	0.786	22.22	2.69	0.917	23.46
0.41	.814	22.70	3.42	.952	23.71
0.87	.832	22.80	6.90	1.108	25.49
1.71	.872	22.99			

Ethyl alcohol (C_2H_6O) + Cadmium chloride ($CdCl_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (3+2)
10	20.5
20	29.6
30	40.5
40	54.0

Ethyl alcohol (C_2H_6O) + Cadmium bromide ($CdBr_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (3+2)
20	40.8
30	50.1
40	60.9

%	f.t.	%	f.t.
21.2	10	28.2	45
23.1	20	27.6	50
25.1	30	23.3	60
27.4	40	18.2	70

Ethyl alcohol (C_2H_6O) + Cadmium iodide (CdI_2)

Ishikawa, Mori and Murooka, 1933

f.t.	%	f.t.	%
0	19.85(2+1)	42	24.24
25	23.83	50	22.00
35	25.32	60	19.61

t	p	
	V + $CdBr_2$ + (2+1)	V + sat.sol. + (2+1)
20	38.78	40.31
22.5	45.08	46.75
25	52.30	54.04
27.5	60.43	62.41
30	69.68	
32.5	80.62	
37	- tr.t.	

Jahn, 1891

%	d	(α)magn. (current=1)
	20°	
25.91	1.54455	120.14
17.15	0.94193	106.16
0	0.79009	84.77

%	n			
	Hα	D	D'	Hβ
		20°		
25.91	1.3950	1.3970	1.3972	1.4027
0	.3591	.3610	-	.3651

Beckmann, 1890

%	D b.t.	%	D b.t.
0	(78.3°)	10.01	+0.398
2.52	+0.092	10.70	.396
2.95	.103	14.55	.610
5.26	.196	16.31	.621
5.91	.207	21.78	1.024

Jones and Getman, 1904

M	D b.t.	M	D b.t.
0.012	+0.018	0.243	+0.305
.041	.050	.252	.316
.042	.052	.312	.392
.058	.068	.327	.420
.069	.080	.471	.597
.089	.110	.478	.610
.119	.142	.529	.630
.122	.142	.786	1.060
.135	.162	.822	.150
.156	.188	1.174	.600
.169	.207	.340	.360
.190	.225	.438	2.100
.220	.272		

Timofeev, 1894

%	f.t.
53.4	17
52.7	23

Muchin , 1913

%	f.t.
33.0	0
50.5	10.8
51.9	11.6
50.8	31.7
50.5	42.1

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
52.5	20	50.8	60
51.9	30	50.8	70
50.9	40	50.9	80
50.7	50		

Doroszewski and Rakovski, 1909

%	d	%	d
	15°		
0	0.79355	30	1.0720
5	.8294	35	.1388
10	.8687	40	.2136
15	.9120	45	.2986
20	.9597	50	-
25	1.0129		

Jahn, 1891

%	d	(α)magn. (current=1)
	20°	
0	0.79009	84.77
19.18	.94730	117.32
32.47	1.09875	150.01

%	n Hα	n D	n D'	n Hβ
		20°		
0	1.3591	1.3610	-	1.3651
19.18	.3834	.3864	1.3859	.3920
32.47	.4059	.4083	.4086	.4152

Getman and Gibbons, 1915

N	%	d	n_D	λ
		20.5°		
0	0	0.7891	1.36104	0.825
0.005	0.116	.7900	.36117	.817
.01	.292	.7911	.36136	.740
.02	.464	.7924	.36148	.684
.05	1.160	.7973	.36219	.650
.1	2.321	.8052	.36355	.644
.2	4.641	.8210	.36559	.639
.25	5.802	.8292	.36695	.646
.5	11.613	.8688	.37289	.640
1.0	23.206	.9484	.38517	-
2.0	46.412	1.1068	.40964	

Helmreich, 1903

%	U 30-0°	U 50-0°	U 70-0°
0	0.57512	0.59310	0.61083
18.60	.48904	.50388	.51712
35.75	.41218	.42381	.43505

Timofeev, 1905

mol%	U
	20°
0	0.5933
0.99	.5655
9.1	.3365

Doroszewski and Rakovski, 1909

%	U	%	U
	15°		
0	0.6552	30	0.4846
5	.6262	35	.4602
10	.5989	40	.4282
15	.5720	45	.4020
20	.5420	50	.3740
25	.5120		

Ethyl alcohol (C_2H_6O) + Cupric chloride ($CuCl_2$)

Etard, 1894

%	f.t.	%	f.t.
32.0	0	38.5	38
35.7	19	41.7	50
35.9	20		

Lloyd, Brown and al., 1928

t	p dissoc. (2+1)
10	19.6
20	31.3
30	43.3
40	62.0

%	f.t.	%	f.t.
29.7	0	36.8	40
31.5	10	39.0	50
33.3	20	41.5	60
35.1	30		

Gladstone and Hibbert, 1897

24.24% 18° mol. refract. = 35.06

Beetz, 1879

$d_{20°}$	Heat Conductivity . 10^5
6-14°	
0.804 (0%)	360
.828	344
.326	326
28-36°	
0.804 (0%)	570
.828	553
.892	526

Ethyl alcohol (C_2H_6O) + Mercuric chloride ($HgCl_2$)

Heterogeneous equilibria.

Beckmann, 1890

%	b.t.	%	b.t.
0	78.3	10.28	78.800
2.84	78.427	14.82	79.052
5.62	78.560	23.13	79.575

Skinner, 1892

%	b.t.
0	78.43
10.0	78.95
19.4	79.38

Timofeev, 1891

mol%	f.t.
7.09	8.5
7.46	20
8.62	38.2

Etard, 1894

%	f.t.	%	f.t.
3.0	-60	31.3	14
7.8	-55	32.0	19
8.8	-43	34.2	31
9.8	-40	36.4	43
14.3	-30	38.9	51
18.6	-23	42.1	62
19.1	-21	42.5	63
21.9	-20	44.7	68
22.1	-17	45.2	75
24.7	-11	48.0	80
27.0	-9	51.0	92
29.7	-5	51.4	93
29.0	0	53.6	100
30.0	+3	60.6	115
30.9	7	65.3	127
31.3	10	67.8	138

Timofeev, 1894

%	f.t.	%	f.t.
21.3	-18	32.2	20
22.2	-17	35.75	35.2
31.0	+ 8.5		

Properties of phases .

Blümcke, 1884

%	t	d
0	15	0.797
3.12	18.2	.822
5.94	18	.845
9.87	18	.877
14.04	18.2	.912
17.96	18.1	.947

Schröder, 1886

%	d	%	d
	20°		
0	0.81290	12.43	0.9103
4.42	0.8448	19.32	0.9736
8.56	0.8774		

Schröder, 1886

t	d	t	d
22.46%		8.56%	
0	1.02855	0	0.89657
17.10	.01100	15.10	.88309
21.00	.00795	26.60	.87256
27.10	.00066		
19.32%		4.42%	
0	0.99513	0	0.86354
18.67	.97686	15.10	.85048
21.0	.97503	18.28	.84764
28.74	.96749	26.60	.84032
		29.95	.83733
15.91%		2.38%	
0	0.96204	0	0.84843
18.10	.94491	18.28	.83287
18.67	.94405	29.95	.82267
28.00	.93525		
28.74	.93500		
12.43%		1.22%	
0	0.93075	0	0.83971
18.10	.91418	18.28	.82428
28.00	.90457	29.95	.81410

Gerlach, 1889

%	d	%	d
	0°		
0	0.80633	12	0.90146
4	.83539	30	1.08647
8	.86663		

Tammann and Hirschberg, 1894

%	d (0°=1)			
	0°	10°	20°	30°
0	1	0.98940	0.97881	0.96803
12.00	1	.98944	.97901	.96844
25.09	1	.98950	.97932	.96886
35.70	1	.98971	.97953	.96903

Schönrock, 1895

%	d
	16°
0	0.79433
11.8801	.88572
23.5489	.99885

P.P. and M.S. Kosakewitsch, 1933

mol%	t	d	mol%	t	d
0.00	18	0.791	0	21	0.789
0.43	18	.817	2.83	21	.919
1.41	18	.850	4.83	21	.995

P.P. and M.S. Kosakewitsch, 1933

mol%	t	σ	mol%	t	σ
0.00	18	22.97	0	21	22.72
0.43	18	23.06	2.83	21	23.31
1.41	18	23.37	4.88	21	23.59

Schönrock, 1895

%	(α)magn.
	16°
0	1.0759
11.8801	0.9314
23.5489	0.8374

Blümcke, 1884

%	t	U	%	t	U
0	15	0.683	9.87	18	0.632
3.12	18.2	.667	14.04	18.2	.612
5.94	18	.652	17.96	18.1	.595

Timofeev, 1891 and 1905

mol%	U
10-50°	
0	0.5933 (20°)
0.99	.57865
1.96	.57683
3.85	.52403

Ethyl alcohol (C_2H_6O) + Mercuric bromide ($HgBr_2$)

Timofeev, 1894

%	f.t.	%	f.t.
20.1	0	24.2	39
20.8	10	30.8	65
23.1	19	40.1	89

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
21.4	0	25.4	40
22.0	10	26.6	50
22.2	20	29.7	60
24.0	30	31.6	70

Ethyl alcohol (C_2H_6O) + Manganese chloride ($MnCl_2$)

Blümcke, 1884

%	t	d	U
5.35	18	0.832	0.667
10.43	18.4	.867	.654
15.12	18.1	.901	.640
21.48	17.5	.948	.626
25.99	18	.988	.615

Quartaroli, 1916

Specific magnetic susceptibility (see author) .

Ethyl alcohol (C_2H_6O) + Cobalt chloride ($CoCl_2$)

Quartaroli, 1916

Specific magnetic susceptibility (see author) .

Ethyl alcohol (C_2H_6O) + Cobalt bromide ($CoBr_2$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
41.4	10	51.2	50
43.5	20	54.7	60
45.9	30	73.0	70
49.0	40	74.6	80

tr.t. = 65°

t	p dissoc.	t	p dissoc.
	(3+1) - (2+1)		
10	1.51	40	25.6
20	4.16	50	50.6
30	10.5		

Ethyl alcohol (C_2H_6O) + Ferric chloride ($FeCl_3$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
57.6	0 (2+1)	65.3	40
58.5	15	66.0	30
59.0	20.6	66.5	20.6 E
59.8	30	67.0	30
60.8	40	67.4	40
63.8	50		

Blümcke, 1884

%	t	d	U
0	15	0.898	0.683
5.12	18	.836	.658
10.00	18	.873	.635
14.30	19	.903	.618
22.00	18	.962	.584
32.10	15	1.062	.538

Gladstone and Hibbert, 1897

%	molecular refraction
12.84	49.32
18.28	49.98
29.73	51.11

Quartaroli, 1916

Specific magnetic susceptibility. (see author)

Beetz, 1879

d 20°	Heat conductivity 6-14°	28-36°
0.804 (0%)	360	570
0.893	322	531

Ethyl alcohol (C_2H_6O) + Stannous chloride ($SnCl_2$)

Cheneveau, 1907

%	d	n_D
	16°	
0	0.7951	1.3639
8.53	.8455	.3731
16.10	.8956	.3834

Ethyl alcohol (C_2H_6O) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altman, 1934

M	κ	M	κ
	0°		
0.0436	0.2530	0.4160	0.6340
.0868	.2980	.4650	.7470
.1293	.3830	.7950	1.0440
.1715	.3820	1.4560	.2800
.2550	.4750	.7500	.1770

Ethyl alcohol (C_2H_6O) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaya, 1937

mol%	d 70°	d 85°	η 70°	η 85°
0	0.7648	-	513	-
10	1.4204	1.3809	2285	1360
18	.8420	.8178	5985	3527
20	.9301	.9072	7151	3935
22	2.0007	.9744	7699	4284
23	.0552	2.0182	7791	4310
24	.0905	.0590	7839	4355
25	.1269	.0924	7712	4260
26	.1325	.1269	7591	4115
33.34	.3420	.3089	5492	3234
50	.6833	.6537	2843	2157
60	-	.8387	-	1732
75	-	3.0058	-	1466
100	3.2488	.1864	1680	1234

Ethyl alcohol (C_2H_6O) + Zinc chloride.4-Picoline ($ZnC_6H_7NCl_2$)

Flaschner, 1909

%	f.t.	sat.t.	%	f.t.	sat.t.
86.9	88.7	-	38.5	68.5	32.9
75.5	78.8	25.5	23.8	64.5	28.4
61.6	73.1	32.4	10.7	55.0	12.8
50.2	71.0	33.4	5.0	43.5	-

Ethyl alcohol (C_2H_6O)+ Uranyl nitrate (UN_2O_8)

Nichols and Merrits, 1914

Fluorescent bands.

Propyl alcohol (C_3H_8O) + Sodium iodide (NaI)

Kosakewitsch, 1928

mol%	d	σ
	24°	
0.00	0.802	23.64
2.32	.843	24.00
4.51	.878	24.32
8.81	.951	24.98

Chatterji and Bose, 1950

t	η(alcohol=1)	t	η(alcohol=1)
16.67%		22.90%	
40	1.605	40	1.978
45	.594	45	.973
50	.575	50	.957
55	.552	55	.938

Propyl alcohol (C_3H_8O) + Magnesium chloride ($MgCl_2$)

Olmer, 1938

%	d
17°	
0	0.805
5	.830
10	.856
15	.881
29	.901

Propyl alcohol (C_3H_8O) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

% (6+1)	f.t.	% (6+1)	f.t.
77.9	0	93.0	43
81.5	10	94.3	46
85.1	20	95.8	48
89.5	30	97.8	50
92.0	40	100	52

Propyl alcohol (C_3H_8O) + Calcium chloride ($CaCl_2$)

Bonnell and Jones, 1926

t	p.dissoc. (3+1)
30	6.42
40	16.31
50	34.19

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
7.6	0 (3+1)	16.6	30 (3+1)
10.6	10	25.3	60
13.6	20		

Kosakewitsch, 1928

mol%	d	σ
	24°	
0.00	0.802	23.64
1.17	.815	29.3
2.32	.829	24.07
4.55	.855	24.28

Propyl alcohol (C_3H_8O) + Calcium bromide ($CaBr_2$)

Bonnell and Jones, 1926

t	p dissoc. (3+1)
30	1.71
40	3.71
50	7.62
60	15.50

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
6.19	0 (3+1)	33.0	40
11.5	10	38.8	50
18.4	20	43.4	60
25.9	30		

Propyl alcohol (C_3H_8O) + Zinc iodide (ZnI_2)

Kosakewitsch, 1928

mol%	d	σ
	24°	
0.00	0.802	23.64
0.83	.827	23.96
1.36	.852	24.02
2.74	.901	24.53
5.50	1.000	25.32

Propyl alcohol (C_3H_8O) + Cadmium iodide (CdI_2)

Timofeev, 1894

%	f.t.
37.1	17
41.6	20
39.4	23

Muchin, 1913

%	f.t.	%	f.t.
15.7	0	34.8	31.8
34.3	13.8	35.1	34
36.3	16.2	35.1	42.1
38.0	21.1		

Propyl alcohol (C_3H_8O) + Mercuric chloride ($HgCl_2$)

Timofeev, 1891

mol%	f.t.
4.69	8.5
5.26	20
6.41	38.2

Timofeev, 1894

%	f.t.	%	f.t.
16.8	-18	20.1	+21.5
17.2	-17	23.6	+36.5
18.15	+8.5		

Etard, 1894

%	f.t.	%	f.t.
14.7	-32	27.9	53
15.4	-22	29.4	62
15.6	-14	32.7	67
16.4	0	36.4	78
16.5	0	43.8	100
18.2	16	52.7	127
23.8	41		

Propyl alcohol (C_3H_8O) + Mercuric bromide ($HgBr_2$)

Timofeev, 1894

%	f.t.	%	f.t.
12.76	0	17.2	39
13.5	10	23.8	65
13.4	19	30.0	86.5

Propyl alcohol (C_3H_8O) + Cupric chloride ($CuCl_2$)

Etard, 1894

%	f.t.	%	f.t.
26.8	-15	30.3	57
30.9	-19	30.5	62
30.7	37		

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
16.7	10	25.5	40
19.8	20	27.4	50
22.7	30	29.1	60

Propyl alcohol (C_3H_8O) + Cobalt chloride ($CoCl_2$)

Lloyd, Brown and al., 1928

t	p dissoc. (2+1)
10	0.85
20	2.21
25	3.67 (3+1)
30	5.54

%	f.t.	%	f.t.
31.0	0	40.3	40
32.7	10	39.4	50
35.2	20	40.0	60
38.5	30	42.2	70
		41.3	80

Propyl alcohol (C_3H_8O) + Zinc chloride.4-Picoline ($ZnC_6H_7NCl_2$)

Flaschner, 1909

%	f.t.	sat.t.	%	f.t.	sat.t.
90.3	96.0	-	26.6	70.2	40.7
78.3	84.0	29.2	17.0	67.1	35.4
65.1	78.6	40.7	10.0	60.5	25.0
50.9	76.3	52.5	5.0	50.0	-
39.7	73.8	42.0			

Isopropyl alcohol (C_3H_8O) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Vold, Leggett and Mac Bain, 1940 (fig.)

%	f.t.
0.6	50
1.8	100

%	clearing point
2	100
10	137
40	160
60	175
80	208
90	233
92	235
100	297

Leggett, Vold and Mac Bain, 1942

%	f.t.	%	f.t.
0.4	30	75.0	200
0.6	50	94.0	250
1.0	90	97.6	280
24.0	150		

Isopropyl alcohol (C_3H_8O) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

% (4+1)	f.t.	% (4+1)	f.t.
40.0	0	62.5	110
42.2	20	67.3	120
45.0	40	74.0	130
48.5	60	83.6	136
53.3	80	90.0	138
59.0	100	100.0	139

Isopropyl alcohol (C_3H_8O) + Magnesium iodide (MgI_2)

Menshutkin, 1907

% (6+1)	f.t.	% (6+1)	f.t.
57.1	10	76.2	110
60.0	30	79.4	120
63.3	50	84.8	130
67.0	70	91.7	136
71.2	90	100	138

Isopropyl alcohol (C_3H_8O) + Cadmium iodide (CdI_2)

Muchin, 1913

%	f.t.
36.9	0
37.3	15.7
37.3	31.2

Isopropyl alcohol (C_3H_8O) + Cupric chloride ($CuCl_2$)

Etard, 1894

%	f.t.
11.0	32
28.3	70
28.7	84

Butyl alcohol ($C_4H_{10}O$) + Magnesium chloride ($MgCl_2$)

Olmer, 1938

%	d
	17°
0	0.810
5	.832
10	.856
15	.879
20	.901

Butyl alcohol ($C_4H_{10}O$) + Calcium chloride ($CaCl_2$)

Bonnell and Jones, 1926

t	p.dissoc. (3+1)
30	1.45
40	3.70
50	9.14
60	13.45

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
13.9	0	24.0	40
17.6	10 (3+1)	25.1	50
20.4	20	25.8	60
22.5	30		

Butyl alcohol ($C_4H_{10}O$) + Calcium bromide ($CaBr_2$)

Bonnell and Jones, 1926

t	p.dissoc. (3+1)
30	1.55
40	2.82
50	5.05
60	9.44

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
15.5	0 (3+1)	35.2	40
20.2	10	39.3	50
25.3	20	42.9	60
30.4	30		

Butyl alcohol ($C_4H_{10}O$) + Mercuric chloride ($HgCl_2$)

Etard, 1894

%	f.t.	%	f.t.
12.4	-21	15.9	21
13.0	-6	25.8	59
14.3	+9	33.1	82

Butyl alcohol ($C_4H_{10}O$) + Cupric chloride ($CuCl_2$)

Etard, 1894

%	f.t.	%	f.t.
15.2	0	16.1	55
15.3	23	16.2	84
15.7	37	16.7	92

Isobutyl alcohol ($C_4H_{10}O$) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

% (6+1)	f.t.	% (6+1)	f.t.
55.8	0	82.4	60
60.5	10	84.2	65
65.2	20	88.0	71
69.3	30	92.0	75
74.3	40	94.6	77
78.5	50	100	80

Isobutyl alcohol ($C_4H_{10}O$) + Zinc chloride.4-Picoline ($ZnC_6H_7NCl_2$)

Flaschner, 1909

%	f.t.	sat.t.	%	f.t.	sat.t.
88.3	93.6	-	37.9	78.4	54.6
82.1	-	35.0	25.9	76.7	52.1
76.3	87.5	44.7	14.9	72.5	45.4
65.5	82.7	52.3	8.0	64.3	35.6
49.5	80.0	55.0	4.0	54.7	22.7

Isobutyl alcohol ($C_4H_{10}O$) + Mercuric chloride ($HgCl_2$)

Etard, 1894

%	f.t.	%	f.t.
5.5	-11	32.1	93
6.2	-6	42.0	127
6.7	0	47.2	145
7.5	11	50.4	155
19.3	63		

Tert.butyl alcohol ($C_4H_{10}O$) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

%(4+1)	f.t.	%(4+1)	f.t.
0.06	24.4	50.5	65
1.0	25	62.5	70
9.5	35	77.0	75
19.1	45	85.0	77.5
32.2	55	91.5	79
40.5	60	100	80

Amyl alcohol ($C_5H_{12}O$) + Sodium Amyl Camphor carbonate ($NaC_{16}H_{25}O_3$)

Brühl and Schröder, 1904

%	t	d
0	14.60	0.8142
7.57	16.80	.8301
13.89	17.20	.8447
19.60	17.00	.8589
24.71	17.45	.8716

%	t	η Hα	η D	η Hβ	η Hγ
0	14.60	1.40747	1.40963	1.41452	1.41870
7.57	16.80	.41332	.41564	.42109	.42558
13.89	17.20	.41888	.42124	.42709	.43204
19.60	17.00	.42425	.42673	.43291	.43824
24.71	17.45	.42913	.43171	.43821	.44391

Amyl alcohol ($C_5H_{12}O$) + Calcium chloride ($CaCl_2$)

Lloyd , Brown and al., 1928

%	f.t.	%	f.t.
6.54	10 (3+1)	22.2	50
10.3	20	25.6	60
14.5	30	29.2	70
18.4	40		

Amyl alcohol ($C_5H_{12}O$) + Calcium bromide ($CaBr_2$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
14.8	10	34.0	50
20.2	20	38.4	60
25.4	30	42.7	70
30.1	40		

Amyl alcohol ($C_5H_{12}O$) + Mercuric chloride ($HgCl_2$)

Etard, 1894

%	f.t.	%	f.t.
8.6	-13	29.8	90
8.9	+26	35.1	106
14.0	50		

Isoamyl alcohol ($C_5H_{12}O$) + Sodium iodide (NaI)

Kosakewitsch, 1928

mol%	d	σ
	22°	
0.00	0.810	24.00
3.81	0.853	24.15
6.26	0.882	24.22

Isoamyl alcohol ($C_5H_{12}O$) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1907

%(6+1)	f.t.	%(6+1)	f.t.
70.2	0	88.7	38
75.6	10	90.0	40
80.2	20	92.0	42
84.5	30	94.2	44
86.7	35	100	46

Isoamyl alcohol ($C_5H_{12}O$) + Calcium chloride ($CaCl_2$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
2.91	10	18.6	50
6.72 (3+1)	20	21.2	60
10.8	30	24.0	70
14.5	40		

Isoamyl alcohol ($C_5H_{12}O$) + Calcium bromide ($CaBr_2$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
17.83	10 (3+1)	30.26	50
20.38	20	33.33	60
23.02	30	35.98	70
26.95	40		

Isoamyl alcohol ($C_5H_{12}O$) + Zinc chloride-4-Picoline ($ZnC_6H_7NCl_2$)

Flaschner, 1909

%	f.t.	sat.t.	%	f.t.	sat.t.
88.1	99.8	-	35.0	83.9	61.8
80.0	93.7	46.5	20.5	79.8	57.1
69.6	88.1	57.9	10.8	74.0	46.5
59.4	86.5	62.2	4.7	61.0	-
50.0	85.4	62.9			

Isoamyl alcohol ($C_5H_{10}O$) + Cupric chloride($CuCl_2$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
9.67	10	14.24	40
10.79	20 (2+1)	17.36	50
12.43	30	20.13	60

Isoamyl alcohol ($C_5H_{10}O$) + Ferric chloride ($FeCl_3$)

Quartaroli, 1916

Specific magnetic susceptibility (see author) .

Heptyl alcohol ($C_7H_{16}O$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggett, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
0.3	50	69.5	200
3.1	90	93.4	250
31.6	150	97.5	280

sec.Octyl alcohol ($C_8H_{18}O$) + Zinc chloride-4-Picoline ($ZnC_6H_7NCl_2$)

Flaschner, 1909

%	f.t.	sat.t.	%	f.t.	sat.t.
86.2	104.0	60.5	38.2	100.0	96.5
73.5	101.8	86.6	22.2	97.7	91.1
60.1	100.3	95.8	10.7	91.7	76.0
49.4	100.2	97.0	3.8	77.0	-

Cetyl alcohol ($C_{16}H_{34}O$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggett, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
1.0	90	85.4	250
5.8	150	95.0	280
44.9	200		

Cetyl alcohol ($C_{16}H_{34}O$) + Zinc chloride.4-Picoline ($ZnC_6H_7NCl_2$)

Flaschner, 1909

%	f.t.	sat.t.	%	f.t.	sat.t.
95.8	108	104	20.0	-	155.0
89.5	-	150	9.7	103.0	120.6
69.7	-	230	4.9	102.5	92.0
50.0	103	241	1.0	54.0	48.0
30.0	-	195			

Octadecyl alcohol ($C_{18}H_{38}O$) + Zinc myristate ($ZnC_{28}H_{54}O_4$)

Martin and Pink, 1948

C.S.T. = 108.0°

Octadecyl alcohol ($C_{18}H_{38}O$) + Zinc stearate ($Zn\ C_{18}H_{35}O_2$)

Martin and Pink, 1948 (fig.)

%	sat.t.	%	sat.t.
0.1	90	1.5	102
0.2	95	4.0	103
0.6	100	17.0	104 C.S.T.

Allyl alcohol (C_3H_6O) + Cupric chloride ($CuCl_2$)

Etard, 1894

	%	f.t.	%	f.t.
23	23.4	-20	22.9	+27
	23.6	+ 4	23.3	+32

Allyl alcohol (C_3H_6O) + Mercuric chloride ($HgCl_2$)

Etard, 1894

%	f.t.	%	f.t.
20.6	-21	35.2	+ 8
29.6	- 1	48.7	+22

Allyl alcohol (C_3H_6O) + Ferric chloride ($FeCl_3$)

Quartaroli, 1916

Specific magnetic susceptibility (see author)

Allyl alcohol (C_3H_6O) + Titanium tetrachloride ($TiCl_4$)

Wertyporoch and Altman, 1934

M	κ	M	κ
	0°		
0.0873	0.9285	0.8060	2.1900
0.1785	1.3720	1.1950	2.0930
0.4210	2.4500	2.2220	1.0550

Allyl alcohol (C_3H_6O) + Stannic chloride ($SnCl_4$)

Wertyporoch and Altman, 1934

M	κ	M	κ
	0°		
0.044	0.0572	0.795	0.5800
0.087	0.0728	1.140	0.6840
0.172	0.1360	1.456	0.7590
0.255	0.2130	1.750	0.8240

Glycol ($C_2H_6O_2$) + Lithium bromide (LiBr)

Gibson and Kincaid, 1937

%	d	%	d
	25°		
1.9962	1.10986	12.360	1.17632
4.069	.12575	20.593	.29017
8.115	.14254	30.100	.38991

Gibson, 1937

%	π (1-1000 atm.)
25°	
0.000	31.23
4.069	30.21
8.115	29.21
12.360	28.24
20.593	26.43
30.100	24.42

Glycol ($C_2H_6O_2$) + Sodium bromide (NaBr)

Gibson and Kincaid, 1937

%	d	%	d
	25°		
0.000	1.10986	9.301	1.18556
1.304	.12003	15.215	.23724
2.077	.12610	19.569	.27833
4.647	.14653	24.983	.33314

Gibson, 1937

%	π (1-1000 atm.)	π (1-1500 atm.)
	25°	
0.000	31.23	16.88
4.647	30.21	16.23
9.361	29.22	15.73
15.215	28.111	15.10
19.569	27.37	14.60
24.983	26.44	13.96

Glycol ($C_2H_6O_2$) + Sodium iodide (NaI)

Gibson and Kincaid, 1937

%	d	%	d
	25°		
0.000	1.10986	15.667	1.24738
3.352	.13702	21.523	.30683
3.807	.14080	30.506	.40917
7.847	.17527	34.879	.46492
10.687	.20058	39.869	.53433

Gibson, 1937

%	π (1-1000 atm.)
25°	
0.000	31.23
3.807	30.66
7.847	30.02
15.667	28.98
21.523	28.36
30.506	27.21
34.879	26.72
39.869	26.12

Glycol ($C_2H_6O_2$) + Sodium bromide (NaBr)

Gibson, 1937

%	π (1-1000atm.)	π (1-500atm.)
	25°	
0.000	31.23	16.88
4.647	30.21	16.23
9.361	29.22	15.73
15.215	28.111	15.10
19.569	27.37	14.60
24.983	26.44	13.96

Glycol ($C_2H_6O_2$) + Potassium bromide (KBr)

Semenchenko and Chakhparonov, 1948

%	f.t.	%	f.t.
12.0	-0.7	13.5	25.7
12.5	+7.7	14.0	31.5
13.0	17.0	15.0	51.8

Glycol ($C_2H_6O_2$) + Potassium iodide (KI)

Gibson and Kincaid, 1937

%	d
25°	
0.000	1.10986
6.857	.16390
13.469	.22064
20.924	.29099
28.526	.37114

Getman, 1908

%	d	η
	25°	
0	1.1097	17330
4.067	.1411	17671
8.421	.1782	18243
12.74	.2140	18644
15.96	.2433	19040

Gibson, 1937

%	π (1-1000 atm.)
25°	
0.000	31.23
6.857	30.35
13.469	29.43
20.924	28.50
28.526	27.58

Glycol ($C_2H_6O_2$) + Cadmium iodide (CdI_2)

Gibson, 1937

%	d	π (1-1000atm.)
0.000	1.10986	31.24
5.501	.15667	30.84
10.102	.20482	30.56
17.567	.28598	30.09
29.650	.44283	29.32
38.461	.58328	28.72
49.830	.80963	27.94

Diethylene glycol ($C_4H_{10}O_3$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Vold, Leggett and Mc Bain, 1940 (fig.)

%	clearing point
2	47
10	91
40	151
70	170
80	182
90	210
95	250
100	297

%	f.t.
1.5	50
13.5	100

Leggett, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
0.4	30	87.2	200
1.5	50	90.0	250
9.2	90	97.3	280
39.9	150		

Glycerol ($C_3H_8O_3$) + Sodium hydroxide (NaOH)

Di Ciommo, 1901

%	$\varkappa \cdot 10^{10}$ 14°	18°	24°	30°
0.38	52.8	72.1	111.0	163.5
1.15	172.2	236.2	329.5	531.5
1.91	240.7	329.5	472.0	708.6
2.61	271.8	386.2	568.7	850.4
3.62	287.0	430.5	637.8	943.9
5.00	297.6	439.9	669.7	1020.5
6.50	265.8	398.6	621.9	943.9
7.48	231.7	345.5	558.1	879.7
8.52	190.3	287.0	490.7	776.0
9.10	157.5	255.1	441.1	720.7
10.31	117.6	191.3	333.9	579.3
13.40	46.3	85.0	166.7	324.6

%	$\tau \cdot 10^3$ 14°	18°	24°
0.38	91	90	79
1.15	93	81	83
1.91	92	72	83
2.61	105	78	82
3.62	113	80	87
5.00	125	87	86
6.50	125	93	96
7.48	123	103	96
8.52	127	118	97
9.10	155	122	106
10.31	157	124	123
13.40	208	160	157

Glycerol ($C_3H_8O_3$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Vold, Leggett and Mc Bain, 1940 (fig.)

%	clearing point	%	clearing point
0	45	60	230
5	74	78	276
10	84	80	275
30	105	85	285
40	115	100	297
50	140		

%	clearing point
1.4	50
24.0	100

Leggett, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
0.3	30	67.4	250 superneat
1.0	50	84.4	" "
15.8	90	85.6	" neat
50.8	150	91.0	280
53.6	200		

Glycerol ($C_3H_8O_3$) + Potassium iodide (IK)

Getman, 1908

%	d	η
	25°	
0.000	1.2474	43973
5.879	.2828	41144
4.335	.3060	39718
11.758	.3168	38274
13.770	.3649	35433

Glycerol ($C_3H_8O_3$) + Potassium hydroxyde (KOH)

Di Ciommo, 1901

%	$\kappa.10^{10}$ 13°	$\tau.10^3$ 13°	$\kappa.10^{10}$ 18°	$\tau.10^3$ 18°
0.47	47.3	90	69.1	90
0.95	89.5	90	133.9	98
1.33	125.9	99	188.5	90
1.80	171.7	99	257.0	89
2.28	195.6	109	303.0	94
3.23	247.7	109	382.7	90
5.32	306.1	113	479.4	99
7.03	325.8	110	504.9	100
8.55	318.9	103	483.7	100
10.26	265.8	128	435.8	100
13.01	180.7	141	308.3	114
14.34	129.7	160	233.9	129
15.58	105.3	158	188.5	120
16.60	72.4	177	136.4	125

%	$\kappa.10^{10}$ 24°	$\tau.10^3$ 24°	$\kappa.10^{10}$ 30°
0.47	116.9	89	163.4
0.95	212.6	70	300.2
1.33	290.6	94	455.6
1.80	384.6	90	605.9
2.28	473.0	80	701.6
3.23	590.0	84	887.6
5.32	766.2	101	1233.1
7.03	807.9	100	1296.9
8.55	776	107	1275.7
10.55	701.6	114	1185.2
13.01	519.8	106	850.4
14.34	414.6	112	693.9
15.58	824.2	114	548.5
16.60	238.6	144	445.2

Glycerol ($C_3H_8O_3$) + Rubidium chloride (RbCl)

Davis and Jones, 1912-13

N	λ 25°	35°	45°	55°	65°	75°
1.00	0.374	0.738	1.320	2.205	3.400	4.994
0.770	.371	.740	.337	.215	.421	5.014
.500	.386	.767	.371	.307	.570	.272
.250	.376	.771	.390	.310	.585	.382
.100	.380	.771	.410	.386	.749	.684
.020	.401	.819	.502	.561	4.003	6.014
.010	.420	.848	.563	.640	.180	.343
.0050	.428	.860	.588	.720	.282	.403
.0025	.432	.887	.633	.784	.408	.542
.00125	.444	.874	.625	.774	.358	.532
.000625	.448	.892	.643	.831	.462	.574

mol%	η (water^t=1) 25°	35°	45°	55°	65°	75°
100	5.240	2.439	1.264	0.6867	0.4177	0.2595
75	.351	.500	.282	.6922	.4195	.2589
50	.542	.547	.307	.7043	.4341	.2604
25	.711	.616	.328	.7130	.4423	.2608
10	.818	.669	.346	.7224	.4470	.2606
0	.880	.683	.350	.7218	.4505	.2612

Glycerol ($C_3H_8O_3$) + Rubidium bromide (RbBr)

Davis and Jones, 1912-13

N	λ 25°	35°	45°	55°	65°	75°
1.00	0.368	0.717	1.272	2.143	3.056	4.595
0.770	.360	.716	.311	.198	.424	.604
0.250	.363	.732	.339	.217	.489	5.120
0.100	.369	.752	.385	.324	.676	.483
0.020	.379	.785	.456	.473	.876	.763
0.010	.409	.835	.483	.623	4.153	6.236
0.0050	.427	.855	.592	.698	.291	.477
0.0025	.451	.879	.627	.711	.348	.592
0.00125	.470	.893	.633	.724	.401	.666
0.000625	.480	.932	.700	.722	.499	.755

Davis and Jones, 1912-13

mol%	η(water^t=1)					
	25°	35°	45°	55°	65°	75°
100	4.965	2.307	1.199	0.6483	0.3967	0.2457
75	5.177	.393	.228	.6603	.4041	.2484
50	.384	.472	.274	.6780	.4108	.2528
25	.623	.565	.284	.6990	.4146	.2554
10	.858	.650	.338	.7096	.4226	.2589
0	.885	.664	.358	.7121	.4259	.2604

Glycerol ($C_3H_8O_3$) + Rubidium iodide (RbI)

Davis and Jones, 1912-13

mol%	η (water^t=1)					
	25°	35°	45°	55°	65°	75°
100	4.613	2.163	1.121	0.6184	0.3763	0.2355
75	4.912	.292	.181	.6436	.3892	.2420
50	5.159	.365	.209	.6585	.3981	.2472
25	.566	.554	.294	.6955	.4132	.2548
10	.769	.628	.328	.7102	.4236	.2566
0	.854	.669	.341	.7144	.4243	.2581

N	λ					
	25°	35°	45°	55°	65°	75°
1.00	0.355	0.704	1.275	2.111	3.229	4.771
0.770	.342	.688	.252	.100	.286	.812
0.500	.334	.680	.236	.089	.278	.829
0.250	.321	.649	.216	.034	.197	.703
0.100	.323	.657	.213	.007	.103	.667
0.020	.345	.703	.330	.256	.575	5.408
0.010	.356	.721	.361	.320	.687	.466
0.0050	.363	.745	.373	.360	.861	.743
0.0025	.368	.751	.391	.408	.937	.832
0.00125	.370	.781	.380	.405	.941	.756
0.000625	.373	.805	.389	.345	4.150	.809

Glycerol ($C_3H_8O_3$) + Mercuric chloride ($HgCl_2$)

Moles and Marquina, 1924

44.22% f.t.= 25°

Benzyl alcohol (C_7H_8O) + Calcium bromide ($CaBr_2$)

Lloyd, Brown and al., 1928

%	f.t.	%	f.t.
11.4(3+1)	10	14.8	50
13.0	20	14.7	60
14.5	30	14.5	70
15.0	40		

Benzyl alcohol (C_7H_8O) + Zinc chloride.4-Picoline ($ZnCl_2.C_6H_7N$)

Flaschner, 1909

%	f.t.	%	f.t.
90.0	101.7	50.4	57.1
76.7	86.7	39.3	46.3
70.1	70.1	26.3	34.0

Erythritol($C_4H_{10}O_4$) + Antimony tribromide ($SbBr_3$)

Pushin and Dezelic, 1932

mol %	f.t.	E
0	118	-
10	109	-
20	98	-
30	88	-
40	78	56.5
50	64	51
56	56.5	56.5
60	60	-
64	63	51
70	68	49
80	76.5	-
90	86	-
100	93	-

Phenol (C_6H_6O) + Silver nitrate ($AgNO_3$)

Bailey, 1930

%	f.t.	%	f.t.
0	41	39.8	1.4
11.9	35.5	41.5	16.0
22.1	27.2	43.2	30.2
31.8	11.4	39.0	-5.3
37.5	0.0 E	47.4	+7.6

(2+1)

Phenol (C_6H_6O) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911-12

%	f.t.	E	%	f.t.	E
0	41	-	65.8	25 (2+1)	5
16.2	35	-	70.6	30	5
25.6	30	-	77.2	35	5
33.2	25	-	83.0	37	-
38.7	20	-	83.7	36.5	-
43.6	15	-	86.6	45	-
48.0	10	-	90.6	55	-
52.0	5	-	92.8	60	-
55.3	10	5	95.2	65	-
58.6	15	5	98.2	70	-
62.1	20	5	100	73	-

Phenol (C_6H_6O) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911-12

%	f.t.	E
0	41	-
13.1	37.5	28.5
22.5	35	28.5
32.3	32.5	28.5
40.0	30	28.5
44.6	28.5	-
49.2	35	28.5
53.0	40 (2+1)	28.5
57.3	45	28.5
62.5	50	28.5
68.5	55	28.5
75.8	60	28.5
84.7	65	-
88.5	65.5	-
91.7	75	66.5
95.8	85	66.5
98.1	90	-
100	94	-

Phenol (C_6H_6O) + Stannic chloride ($SnCl_4$)

Usanovich and Pichugina, 1956

mol%	d 20°	d 40°	d 60°	d 80°
100	2.2340	2.1819	2.1288	2.0790
86.59	1.9882	1.9447	1.8986	1.8650
79.80	.8788	.8354	.8060	.7618
73.89	.7967	.7711	.7256	.6927
67.99	.7250	.6633	.6480	.6070
58.41	.6076	.5879	.5540	.5159
53.90	.5697	.5320	.4951	.4745
51.15	.5346	.4954	.4640	.4340
47.16	.4902	.4552	.4165	.3822
41.17	.4411	.4117	.3802	.3407
33.67	.3679	.3302	.3018	.2708
23.69	.2761	.2433	.2239	.1900
22.78	.2672	.2391	.2151	.1800
0	-	.0601	.0443	.0230

mol%	η 20°	η 40°	η 60°	η 80°
100	927	758	637	540
69.80	1820	1210	864	651
58.78	2800	1630	1080	767
50.54	4170	2190	1380	937
43.39	5970	2730	1550	1030
33.63	11900	4280	2110	1260
29.68	16300	4970	2290	1330
27.42	21600	5560	2410	1360
23.94	25900	6280	2570	1400
20.17	38400	7920	2960	1580
15.49	47400	8640	3150	1640
10.06	42700	8650	3260	1700
9.61	40400	8540	3270	1730
0	-	4790	2480	1520

mol%	ϰ	mol%	ϰ
20°			
58.78	0.057	19.70	0.259
52.00	.128	16.40	.250
44.54	.187	14.22	.245
38.02	.249	12.22	.244
32.16	.293	12.10	.239
27.99	.304	9.49	.225
24.98	.294	8.59	.229
40°			
57.47	0.0611	22.80	0.600
50.13	.129	20.80	.612
42.14	.257	19.19	.615
28.55	.481	16.07	.601
27.10	.539	14.53	.590
24.39	.576	12.74	.565
60°			
41.67	0.192	9.47	0.464
33.08	.342	6.65	.355
26.76	.469	4.19	.231
20.79	.556	2.50	.143
15.14	.579	1.94	.0991
12.34	.536		
80°			
57.26	0.0294	18.47	0.360
38.31	.175	14.49	.339
23.02	.341		

Usanovich and Pichugina, 1956

mol %	p 40°	p 60°	p 80°
100	47	120	259
80.10	46	111	241
59.97	43	103	221
39.97	35	84	175
18.85	21	48	100
10.03	-	26	48

Phenol (C_6H_6O) + Stannic bromide ($SnBr_4$)

Kurnakov and Voskresenskaja, 1937

mol %	d 40°	d 60°	d 80°
0	1.0579	1.0411	1.0231
10	1.3810	1.3601	1.3375
20	1.6756	1.6429	1.6146
30	1.9340	1.8983	1.8652
40	2.1812	2.1408	2.1029
50	2.4069	2.3620	2.3200
60	2.6134	2.5637	2.5166
70	2.8036	2.7511	2.6994
80	2.9888	2.9340	2.8836
90	3.1529	3.0946	3.0405
100	3.3127	3.2511	3.1979

mol %	η 40°	η 60°	η 80°
0	4561	2410	1448
10	4210	2270	1409
20	3817	2148	1375
30	3351	1987	1321
40	3034	1888	1307
50	2762	1791	1277
60	2560	1727	1264
70	2387	1680	1252
80	2262	1653	1271
90	2137	1602	1260
100	2101	1621	1302

o-Cresol (C_7H_8O) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggett, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
0.2	50	65.4	200
1.2	90	91.0	250
9.2	150	95.0	280

m-Cresol (C_7H_8O) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggett, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
1.0	30	73.0	200
2.7	50	93.0	250
10.0	90	97.0	280
52.0	150		

p-Cresol (C_7H_8O) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggett, Vold and Mc Bain, 1942

%	f.t.	%	f.t.
1.8	50	76.4	200
9.0	90	94.0	250
60.2	150	97.6	280

o-Nitrophenol ($C_6H_5O_3N$) + Stannic chloride ($SnCl_4$)

Usanovich and Pichugina, 1956

mol%	d 25°	d 40°	d 60°
100	2.2190	2.1819	2.1233
88.40	2.0731	2.0370	1.9840
64.50	1.8050	1.7620	1.7243
43.61	-	1.5930	1.5500
9.56	-	1.3610	-
8.76	-	-	1.3265
0	-	1.3000	1.2750

mol%	η 25°	η 40°	η 60°
100	881	758	637
80.14	1060	881	719
49.18	1750	1320	980
29.18	-	1820	1250
5.31	-	2620	-
4.79	-	-	1780
0	-	3020	1960

p-Nitrophenol ($C_6H_5O_3N$) + Stannic chloride ($SnCl_4$)

Usanovich and Pichugina, 1956

mol%	m.t.	f.t.	E	
97.4	66.0	-	-	
96.3	77.0	-	-	
93.1	88.0	95.2	-	
89.9	88.5	102.0	-	
88.3	88.0	-	-	
75.2	87.0	110.0	-	
63.4	87.5	111.0	-	
54.1	88.5	108.0	-	(2+1)
51.3	87.5	-	-	
47.1	89.0	104.0	-	
40.4	89.0	89.0	87.0	
38.9	91.0	-	86.0	
36.2	90.5	-	86.0	
34.9	91.5	-	86.0	
29.3	95.5	-	87.0	
25.6	93.5	-	86.5	
23.6	99.0	-	87.0	
19.8	100.5	-	86.0	
16.6	104.5	-	84.0	
7.1	109.0	-	85.0	
0	113.5	-	-	

2,4-Dinitrophenol ($C_6H_4O_5N_2$) + Stannic chloride ($SnCl_4$)

Usanovich and Pichugina , 1956

mol%	f.t.	mol%	f.t.
98	35	60	88
95	45	35	100
90	58	20	105
70	80	0	110

Formic acid (CH_2O_2) + Lithium chloride (LiCl)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	14°	
0	1.226	37.21
0.50	.227	37.83
2.17	.236	38.68
4.62	.248	40.05
7.12	.259	41.23

Formic acid (CH_2O_2) + Lithium bromide (LiBr)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	13-14°	
0	1.224	38.25
1.44	.250	39.17
2.99	.273	40.15
5.95	.318	41.82
10.23	.381	44.08

Formic acid (CH_2O_2) + Lithium iodide (LiI)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	13-14°	
0	1.224	38.25
0.99	.249	39.03
1.65	.259	39.04
2.61	.283	40.09
3.01	.287	40.22
5.94	.344	41.87
6.20	.349	42.28

Formic acid (CH_2O_2) + Lithium formate ($LiCHO_2$)

Groschuff, 1903

%	f.t.	%	f.t.
25.4	0	26.9	60
25.9	18	27.8	79
26.4	39		

Kendall and Adler, 1921

mol%	f.t.	mol%	f.t.
0	8.40	23.49	-23.9
1.58	7.0	24.33	-25.0
3.47	5.2	23.49	+18.0
5.33	3.2	23.93	34.0
7.09	1.1	25.31	80.0
8.93	-1.3	25.91	90.5
10.75	-3.5	26.38	97.8
12.23	-5.6	27.71	113.1
13.99	-8.2	29.87	131.2
18.99	-14.6	31.98	145.1
19.56	-17.1	33.04	150.4
21.25	-19.8	35.01	159.1
22.24	-21.7	36.13	163.5

Formic acid (CH_2O_2) + Sodium chloride (NaCl)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	13-14°	
0	1.224	38.25
0.78	.234	38.80
2.93	.252	39.73

Formic acid (CH_2O_2) + Sodium bromide (NaBr)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	14-15°	
0.00	1.224	38.25
1.93	.261	39.17
3.85	.298	40.19
5.75	.333	41.02

Formic acid (CH_2O_2) + Sodium iodide (NaI)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	13°	
0.00	1.224	38.25
1.90	.281	39.42
4.12	.342	40.70
7.25	.425	42.61

Formic acid (CH_2O_2) + Sodium formate ($NaCHO_2$)

Groschuff, 1903

%	f.t.	%	f.t.
22.35(1+1)	0	41.27	70 (0+1)
29.62	25.5	41.60	73
38.85	45.5	43.09	85
41.08	66.5		

Kendall and Adler, 1921

mol %	f.t.	mol %	f.t.
0 (2+1)	+ 8.4	21.86	26.3
1.04	+ 7.5	23.23	30.5
2.92	+ 5.3	24.43	33.1
4.61	+ 3.0	24.92	34.5
6.35	+ 0.4	25.18	31.0
8.72	- 3.8	25.90	37.3
10.50	- 7.6	27.00	45.2
12.58	-12.8	27.10	45.6
14.52	-17.4	28.39	51.2
16.45	+ 0.3	29.80	59.0
18.18	+10.7	31.15	65.1
20.69	+22.9	31.65	67.5
21.13	+24.4	32.14	69.6
29.80(1+1)	45.1	39.10	118.6
31.65	63.6	42.79	135.2
33.40	81.2	43.47	137.7
35.71	99.3	100	255

Formic acid (CH_2O_2) + Sodium acetate ($NaC_2H_3O_2$)

Davidson and Ramskill, 1941

mol %	f.t.	mol %	f.t.
0	+8.4	7.01	- 6.0
0.76	+7.3	8.56	- 8.3
1.43	+6.0	9.73	-12.0
2.39	+4.5	10.71	-14.5
3.20	+3.1	11.85	-18.2
4.17	+0.9	12.42	-20.3
5.19	-1.5	13.11	-22.5
6.06	-3.7	13.51	-24.0

Formic acid (CH_2O_2) + Potassium chloride (KCl)

P.P. and M.S. Kosakewitsch, 1933

mol %	d	σ
	13-14°	
0.00	1.224	38.25
0.88	1.237	39.03
2.22	1.251	39.61
4.49	1.276	40.62
7.70	1.309	41.78

Formic acid (CH_2O_2) + Potassium bromide (KBr)

P.P. and M.S. Kosakewitsch, 1933

mol %	d	σ
	13-14°	
0.00	1.224	38.25
1.14	1.252	39.00
2.71	1.285	39.66
5.09	1.336	40.98

Formic acid (CH_2O_2) + Potassium iodide (KI)

P.P. and M.S. Kosakewitsch, 1939

mol%	d	σ
	12-13°	
0.00	1.224	38.25
1.17	.264	39.06
1.37	.265	39.39
2.89	.326	40.19
3.48	.334	40.33
5.07	.381	41.13
5.97	.405	41.86

Formic acid (CH_2O_2) + Potassium formate ($KCHO_2$)

Groschuff, 1903

%	f.t.	%	f.t.
36.3	0	44.0	60
38.2	19.5	45.9	70
40.8	39.5	52.1	90

Kendall and Adler, 1931

mol%	f.t.	mol%	f.t.	
0	+8.4	31.97	72.9	
0.97	+7.4	33.74	80.6	
3.02	+4.9	37.29	93.0	
4.73	+2.2	38.34	96.1	
6.36	-0.9	41.68	103.2	
8.43	-5.7	42.63	104.3	
10.74	-12.6	42.91	104.6	
12.63	-18.7	45.95	107.5	
13.98	-23.8	50.25	108.6	(1+1)
15.57	-31.5 E	51.49	108.2	
16.62	-27.3	54.24	107.2	
17.44	-23.5	58.47	103.4	
18.10	-21.7	61.14	100.6	
18.88	-19.5	63.14	98.7	E
19.48	-19.10 (2+1)	66.45	108.1	
19.91	-16.0	68.71	114.5	
21.21	-8.0	71.24	122.3	
22.79	-0.6	75.18	130.7	
23.04	-10.1	77.75	135.8	
24.14	+7.3	82.41	143.6	
25.98	29.9	86.68	150.0	
28.56	53.0	91.24	158.3	
30.41	65.1	100	167.5	

Formic acid (CH_2O_2) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1906-1907

%(6+1)	f.t.	%(6+1)	f.t.
49.8	0	73.1	70
57.5	20	86.0	80
65.1	40	95.0	86
73.1	60	100.0	88

Formic acid (CH_2O_2) + Calcium chloride ($CaCl_2$)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	14°	
0.00	1.226	37.2
0.30	.235	38.3
1.40	.255	38.39
2.27	.269	39.36
2.50	.276	39.58
4.06	.306	40.51

Formic acid (CH_2O_2) + Barium formate ($BaC_2H_2O_4$)

Kendall and Adler, 1921

mol%	f.t.	mol%	f.t.
0	8.4	6.95	-4.9
0.91	7.2	8.52	+9.5
1.74	6.1	8.86	15.5
2.30	5.1	9.23	19.0
3.73	2.6	9.83	24.9
4.67	0.5	10.03	26.5
5.12	-0.3	10.75	31.8
(1+1)			

Acetic acid ($C_2H_4O_2$) + Lithium nitrate ($LiNO_3$)

Davidson and Geer, 1938

mol%	f.t.	mol%	f.t.
0	16.60	9.38	39.0
0.156	16.30	10.31	46.3
0.983	16.03	11.26	54.2
2.013	15.71	12.64	66.7
4.115	15.00	14.66	82.6
5.305	14.46	15.66	91.6
7.693	12.80	17.08	101.0
9.384	11.80	19.84	118.3
10.310	10.92	22.34	132.2

Acetic acid ($C_2H_4O_2$) + Lithium acetate ($LiC_2H_3O_2$)

Lescoeur, 1893

t	p dissoc. (1+1)	p sat.sol.
40	-	30
60	-	82
80	35	145
100	150	250

Davidson and Allister, 1930

%	f.t.	%	f.t.
0.0	16.50	38.48	108.0 (1+1)
3.56	14.85	39.83	109.0 "
3.81	14.80	41.84	110.0 "
5.01	14.16	44.41	111.0 "
5.72	13.75	45.32	111.5 "
6.86	13.15	45.74	112.0 "
7.57	12.90	46.77	112.0 "
7.70	12.75	48.76	112.5 "
8.45	12.10	50.06	112.5 "
8.61	16.8 (1+1)		
9.15	22.1 "	43.95	136.0
9.53	25.0 "	44.92	140.0
11.20	41.5 "	46.04	147.5
12.28	51.0 "	47.57	156.0
13.07	55.6 "	48.22	161.0
14.42	63.2 "	49.22	165.0
16.24	70.0 "	52.41	178.5
18.31	78.0 "	55.39	188.0
20.25	83.0 "	56.90	192.0
29.82	91.0 "	70.87	221.0
27.64	93.0 "	100.0	272.0
30.41	101.3 "		
35.34	106.5 "		

Vasiliev, 1909

mol %	f.t.
9.11	14.2 E
50.0	99 (1+1)

Acetic acid ($C_2H_4O_2$) + Sodium formate ($NaCHO_2$)

Davidson and Ramskill, 1941

mol%	f.t.	mol%	f.t.
0	16.60	22.36	63.1
0.54	16.15	26.34	66.5
1.33	15.60	28.32	68.2
2.96	14.25	30.42	69.0
3.74	13.50	32.17	69.2
4.78	12.75	32.80	69.2
5.79	11.80	33.09	80.0
7.27	10.60	33.51	83.0
8.14	9.70	33.98	87.0
8.75	14.55(2+1)	35.24	109
9.24	17.70	35.73	114
11.08	27.10	46.46	171
12.22	32.25	48.81	173
14.61	40.10	50.17	179
15.73	44.35	60.48	215
18.18	53.0	68.93	227
19.76	56.2	100.00	253

Acetic acid ($C_2H_4O_2$) + Sodium acetate ($NaC_2H_3O_2$)

Lescoeur, 1893

t	p dissoc. (1+1)	p dissoc. (2+1)	p sat.sol.
40	-	-	46
50	-	-	57
60	-	-	75
70	-	-	98
75	110	-	110
80	-	-	125
100	-	-	207.5
110	-	-	270
120	-	62	-
130	-	95	-
140	-	130	-
150	-	220	-

Vasilev, 1909

mol%	f.t.	
5.7	12.1	E
33.3	-	(2+1)

Kendall and Adler, 1921

mol%	f.t.	mol%	f.t.	
0	16.5	33.03	96.25	
0.83	16.1	33.16	96.3	
3.59	14.3	34.03	112.0	(2+1)
5.40	13.1	36.57	132.3	
7.11	25.3	39.06	145.2	
8.92	36.7	43.54	157.0	
12.17	54.3	44.25	160.6	
15.27	66.9	46.28	162.3	
16.58	71.9	48.76	174.4	
21.55	85.7	49.49	195.5	
26.86	93.2	100	-	
30.72	96.1			

Bakunin and Vitale, 1935

mol%	f.t.	E	
0.00	.16	-	
7.51	9	-	
14.99	55	-	
23.46	81	-	
33.01	92	92	(2+1)
41.10	128.5	92	

Acetic acid ($C_2H_4O_2$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Leggett, Vold and Mc Bain, 1942

%	f.t.
75.5	90
81.2	150
85.1	200
95.3	250
98.5	280

Acetic acid ($C_2H_4O_2$) + Potassium acetate ($KC_2H_3O_2$)

Lescoeur, 1893

t	p dissoc. (2+1)	p dissoc. (1+1)	p sat.sol.
40	-	-	20
60	12	-	40
80	32	-	56
100	70	-	70
110	-	-	90
120	-	-	105
130	-	-	115
140	-	-	115
160	-	60	60
170	-	100	100
180	-	130	-

Vasilev, 1909

mol%	f.t.	
10.4	7.9	E
33.3	112	(2+1)

Davidson and Mc Allister, 1930

%	f.t.		%	f.t.	
0.0	16.50		28.44	99.0	(1+1)
1.22	15.82		28.69	101.2	
2.11	15.30		30.67	110.0	
6.70	10.95		31.45	114.0	
7.62	9.71		32.47	118.0	
9.10	7.45		33.44	121.0	
10.03	5.95		33.94	121.5	
			34.41	124.5	
10.97	14.76	(2+1)	34.65	125.5	
11.36	17.95		35.00	126.0	
12.71	28.03		39.50	137.0	
13.30	32.02		44.88	145.0	
15.75	49.90		45.62	146.0	
15.94	51.50		48.30	147.5	
17.00	55.50		50.22	148.0	
17.92	60.1		52.32	147.5	
18.71	64.1				
19.00	65.9		58.45	170	
20.48	73.5		61.40	191	
23.72	83.1		64.16	206	
23.85	83.6		65.21	207	
			70.05	230	
			76.50	249	
			100.0	292	

Bakunin and Vitale, 1935

mol%	f.t.	E	
0.00	16	-	
5.27	9	0.0	
10.38	3	0.0	
16.22	14	0.0	(2+1)
23.43	64	-	(1+2)
29.43	101	63	
37.27	112	60	
42.88	134	56	
50.06	138.6	59	
58.05	147	136	
60.20	175	143	
71.83	234	142	
81.73	258	136	
98.24	292	-	
100.0	297	-	

Acetic acid ($C_2H_4O_2$) + Magnesium bromide ($MgBr_2$)

Menshutkin, 1906-1907

%(6+1)	f.t.	%(6+1)	f.t.
0.3	17	49.5	85
1.5	30	57.7	90
4.5	50	71.8	100
7.9	60	80.0	105
16.2	70	89.5	110
38.5	80	100	112

Acetic acid ($C_2H_4O_2$) + Magnesium iodide (MgI_2)

Menshutkin, 1906-1907

%(6+1)	f.t.	%(6+1)	f.t.
0.6	20	42.0	95
2.0	40	54.5	105
5.0	60	65.0	115
9.5	70	73.8	125
13.0	75	85.0	135
18.5	80	94.0	140
27.1	85	100	142

Acetic acid ($C_2H_4O_2$) + Calcium chloride ($CaCl_2$)

Menshutkin, 1906-07

%(6+1)	f.t.	%(6+1)	f.t.
0	+16.2	54.7	40
18.0	15	63.0	45
27.0	14	69.5	50
34.0	13	79.5	60
42.0 (4+1)	11.1 E	84.5	65
47.6	30	91.2	70
50.0	35	100	73

Menshutkin, 1907

%(4+1)	f.t.	%(4+1)	f.t.
42.0	11.1	69.5	50
44.7	20	75.0	55
47.6	30	79.5	60
50.0	35	84.5	65
54.7	40	91.2	70
63.0	45	100	73

Morgan and Benton, 1907

%	D.f.t.	%	D.f.t.
1.03	-0.63	11.53	-4.99
1.37	0.84	14.12	5.74
1.94	1.20	18.15	6.63
2.80	1.46	25.01	8.67

P.P. and M.S. Kosakewitsch, 1932

mol%	d	σ
	21°	
0	1.048	27.57
0.80	.057	27.31
2.34	.082	28.15
4.42	.111	28.91
7.44	.154	30.79

Acetic acid ($C_2H_4O_2$) + Calcium nitrate (CaN_2O_6)

Davidson and Geer, 1938

mol%	f.t.	mol%	f.t.
0.00	16.60	8.90	14.37
2.70	15.86	7.87	30.3 (3+1)
3.65	15.72	8.27	32.0
4.59	15.10	8.71	33.5
6.57	15.18	8.90	33.9
7.87	14.83		

Acetic acid ($C_2H_4O_2$) + Strontium chloride ($SrCl_2$)

Davidson and Chappell, 1938

mol%	f.t.	mol%	f.t.
0	16.60	1.67	98.9
0.55	16.32	2.22	83.0
1.09	16.05	3.28	60.0
1.81	15.08	4.04	45.0
3.07	15.05	5.12	30.0
3.93	14.50	6.11	19.0
5.11	13.65		

Acetic acid ($C_2H_4O_2$) + Barium acetate ($BaC_4H_6O_4$)

Davidson and Allister, 1930

%	f.t.	%	f.t.
0.0	16.50	6.82	32.0 (2+1)
1.27	15.60	7.53	40.8
1.59	15.43	8.11	48.0
1.91	15.30	8.53	51.5
2.32	15.00	8.62	54.0
3.04	14.2	8.74	55.0
4.77	13.4	9.84	64.0
5.55	12.5	11.37	72.3
6.34	11.4	11.70	74.1
7.47	10.5	12.46	77.9
7.76	9.4	14.06	84.0
		15.08	88.0
2.67	19.0 (3+1)	15.58	90.3
3.16	22.5		
4.48	31.8	8.11	13.8 (2+1)
5.05	35.2	9.70	19.9
5.18	36.0	10.78	24.3 unstable
6.21	41.1	11.37	27.1
6.86	44.0	11.88	29.8
7.53	47.3	12.38	32.0
8.11	48.0		
8.62	49.0		

Acetic acid ($C_2H_4O_2$) + Lead acetate ($PbC_4H_6O_4$)

Davidson and Allister, 1930

%	f.t.	%	f.t.
0.0	16.50	35.69	49.0
0.54	16.33	38.27	68.0
5.18	15.52	38.46	70.2
8.42	14.24	41.43	87.3
11.53	13.40	41.75	88.0
19.78	9.55	44.95	100.5
21.99	8.05	52.08	119
32.97	-0.3	59.55	139
33.52	-0.8	71.83	165
34.21	-1.2	90.20	192
37.48	-3.1	100.0	204

Davidson and Chappell, 1933

mol%	f.t.	mol%	f.t.
13.0	16.62	24.77	40.5
13.06	16.4	27.60	45.2
14.67	20.3	30.23	48.3
16.55	24.5	33.55	52.9
17.37	26.4	36.10	56.4
20.57	32.8		
22.60	36.7		

Acetic acid ($C_2H_4O_2$) + Zinc chloride ($ZnCl_2$)

Davidson and Chappell, 1933

mol%	f.t.	mol%	f.t.
0	16.60	12.88	8.55
1.77	15.75	14.14	7.30
3.46	14.95	16.16	4.80
4.31	14.45	17.80	2.8
7.31	12.85	18.45	2.0
9.67	11.45	22.07	-4.0
10.99	10.35		

Acetic acid ($C_2H_4O_2$) + Zinc bromide ($ZnBr_2$)

Sakhanov, 1910

M	λ
25°	
0.05353	0.014
.105	.018
.225	.037
.532	.12
.800	.19

Acetic acid ($C_2H_4O_2$) + Mercuric chloride ($HgCl_2$)

Etard, 1894

%	f.t.	%	f.t.
2.7	21	12.4	95
3.0	22	16.0	115
5.0	33	17.0	116
6.0	43	20.0	127
6.7	50	26.3	145
8.0	61	44.8	182
11.0	87	55.2	207

Davidson and Chappell, 1938

mol%	f.t.	mol%	f.t.
0	16.60	1.46	53
0.270	16.45	.61	60
.477	16.35	.81	66
.477	17	(2+1) 2.01	72
.680	24	.22	78
.936	30	.42	84
1.06	33	.74	92
.19	36	3.05	97
.15	40	.55	106
.30	46	4.21	117

Acetic acid ($C_2H_4O_2$) + Mercuric bromide ($HgBr_2$)

Davidson and Chappell, 1938

mol%	f.t.	mol%	f.t.
0.194	26.50	0.707	75
.261	25	.806	86
.287	32	.998	92
.350	41	1.13	97
.413	51	.29	103
.477	58	.50	110
.650	71		

Acetic acid ($C_2H_4O_2$) + Nickel acetate ($NiC_4H_6O_4$)

Davidson and Chappell, 1933

mol%	f.t.	mol%	f.t.
0.0	16.55	5.68	15.35
0.32	16.45	7.32	15.10
0.61	16.40	8.90	14.90
1.00	16.30	10.57	14.55
1.73	16.15	12.19	14.15
2.47	16.00	13.67	13.80
3.22	15.85	14.89	13.55
4.47	15.60		

Acetic acid ($C_2H_4O_2$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E
0.0	16.5	-
6.5	15	-
22.7	10	-
34.3	5	-
42.5	0	-
48.5	-5	-9 (1+1)
52.7	-9	- "
55.6	-5	-9 "
59.0	0	-9 "
63.0	5	-9 "
67.3	10	-9 "
72.6	15	-9 "
76.5	18	- "
79.1	19	- "
81.4	18	- "
84.5	15	- "
76.7	5	-
79.1	15	-
81.5	25	-
84.2	35	-
87.4	45	-
91.0	55	-
95.3	65	-
98.0	70	-
100.0	73	-

Usanovich and Sumarokova, 1951

%	mol%	d 20°	d 50°	d 60°
100.00	100.00	crist.	2.7352	2.7148
97.40	91.18	"	.6391	-
95.28	84.23	"	.5667	.5418
93.79	79.90	2.5824	.5174	.4912
91.52	74.06	-	.4644	.4445
90.36	71.23	.4712	.4026	.3831
89.32	68.75	.4356	.3741	.3468
80.64	52.22	.1757	.1176	.1004
80.04	51.32	.1632	.1050	.0914
75.00	44.09	.0353	1.9833	1.9686
53.21	23.07	1.5769	.5693	-
34.75	11.92	.3400	-	-

mol%	η 20°	η 50°	η 60°
100.00	-	3930	3150
91.18	-	4300	-
84.23	-	4770	3600
79.90	19130	5440	4680
74.06	-	5770	-
71.23	23980	5770	4350
68.75	24730	5850	4180
60.21	28030	6280	4410
53.94	29950	6320	4390
52.22	29170	6280	4390
51.32	26760	5880	4160
44.09	23260	4830	3860
38.39	13180	4010	2950
23.07	28600	2840	-
11.92	3570	-	-

mol%	κ 20°	κ 50°	κ 60°
91.18	-	33.56	37.44
84.23	-	51.64	59.40
79.90	26.33	61.28	-
74.06	-	57.23	-
71.23	-	-	-
68.75	22.42	60.40	69.11
60.21	19.26	52.16	62.36
53.94	15.37	45.61	56.85
52.22	14.69	-	-
51.32	14.34	42.72	53.70
44.09	11.11	28.90	38.64
38.39	6.80	17.23	21.29
23.07	5.47	-	-
11.92	1.92	4.23	5.21

Acetic acid ($C_2H_4O_2$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	%	f.t.	E
0	16.5	-	72.5	40	4
12.2	15.0	4	77.0	50	4
41.3	10	4	81.9	60	4
58.2	4	-	87.1	70	4
60.6	10	4	92.4	80	4
64.3	20	4	97.8	90	4
68.3	30	4	100	94	-

Acetic acid ($C_2H_4O_2$) + Ferric chloride ($FeCl_3$)

Quartaroli, 1916

Specific magnetic susceptibility. (see author)

Acetic acid ($C_2H_4O_2$) + Stannic chloride ($SnCl_4$)

Usanovich and Kalabanovskaya, 1947

mol%	0° d	0° η	25° d	25° η	50° d	50° η
0.00	-	-	1.0439	1100	1.0162	750
3.03	1.1795	-	.1507	2500	.1222	1160
4.41	.2230	-	.1957	3390	.1664	1390
5.42	.2538	-	.2250	4369	.1972	2290
6.91	.3028	20330	.2728	6420	.2446	3170
9.78	.3621	46870	.3340	11000	.3021	4470
13.85	.5061	362510	.4750	40560	.4412	11140
14.40	.5239	469250	.4929	48490	.4611	12520
16.73	.5853	1093020	.5526	77070	.5194	13120
22.84	.7296	5673300	.6940	243700	.6569	37340
24.99	-	7293300	-	278520	-	39510
27.20	.8213	7001200	.7794	266100	.7374	36850
33.24	.9068	2099600	.8639	135540	.8147	25420
35.91	.9472	1007700	.8883	69870	.8408	16210

mol%	κ 0°	κ 25°	κ 50°
3.03	7.59	17.02	28.83
4.41	-	22.52	40.72
5.42	-	25.31	47.55
6.91	9.06	26.21	52.59
9.78	6.72	23.37	57.26
13.85	2.19	13.31	40.08
14.40	1.84	12.40	38.86
16.73	1.02	9.05	31.55
22.84	0.26	3.56	17.28
24.99	0.18	2.86	11.67
27.20	0.15	2.51	13.59
33.24	0.25	3.14	13.48
35.91	0.39	3.85	14.49

Stranathan and Strong, 1927

mol%	d	η	κ
		25.2°	
0.00	1.0542	1155	-
2.44	.1311	2121	11.85
5.20	.2352	4373	25.29
10.84	.4071	18970	20.74
17.19	.5798	96300	8.68
22.90	.7153	258200	3.670
25.27	.7683	303400	2.887
26.82	.7921	290900	2.510
29.69	.8362	210500	2.591
31.51	.8653	164900	2.857
34.98	.9048	86400	3.678
36.08	.9209	69100	3.967
41.11	.9410	42340	4.550
100.00	2.2543	840	-

Acetic acid ($C_2H_4O_2$) + Stannic bromide ($SnBr_4$)

Usanovich and Yakovleva, 1955 (fig.)

mol%	d	η		
	25°	25°	60°	85°
0	3.40	2500	1700	1450
20	3.15	2400	1500	1250
30	3.05	2410	1450	1200
50	2.70	-	1400	1050
70	2.20	-	1450	1100
80	1.95	5500	1600	1150
88	-	6450	1600	1150
100	1.10	0	0	0

mol%	κ		
	25°	60°	75°
50	0.1	0.3	20.1
60	0.5	1.5	0.5
70	3.5	5.0	3.5
80	13	20	13
90	29.5	39.5	32
100	0	0	0

Butyric acid ($C_4H_8O_2$) + Sodium butyrate ($NaC_4H_7O_2$)

Bakunin and Vitale, 1935

mol%	f.t.	E	
0.00	-2	-	
7.11	-6	-	
14.56	-13	-27	
22.12	-22	-	
29.76	20	-26	(1+1)
39.51	83.1	-	
47.22	127	48	
53.14	155	45	
67.80	196	46	
72.93	206	42	
84.12	225	45	
92.90	238	42	
100	286	-	

Butyric acid ($C_4H_8O_2$) + Calcium chloride ($CaCl_2$)

P.P. and M.S. Kosakewitsch, 1933

mol%	d	σ
	13-16.5°	
0	0.961	26.57
0.51	.9617	26.94
1.15	.970	27.10
2.40	.980	27.39
3.70	.998	27.79

Valeric acid ($C_5H_{10}O_2$) + Stannic chloride ($SnCl_4$)

Bingham, 1907

C.S.T. = -10°

Lauric acid ($C_{12}H_{24}O_2$) + Potassium laurate ($KC_{12}H_{23}O_2$)

Mc Bain and Field, 1933

mol%	f.t.	mol%	f.t.
0	44.0	33.10	85.5
1.06	43.7	36.04	89.7
2.91	43.3	38.77	93.0
4.03	42.7	41.23	107.1
7.00	41.0	42.03	115.0
9.19	43.5	43.92	125.0
13.30	52.6	46.55	144.7
16.63	57.0	49.51	160.0
19.72	63.9	53.86	185.0
22.66	69.0	56.94	200.5
22.69	69.2	61.70	211.5
24.92	74.3	64.19	217.5
28.18	80.0		

mol%	f.t.	sat.t.
66.58	226.0	-
71.55	235.5	259.5
75.15	-	305.0
79.87	-	358.0
85.69	-	374.0
92.34	-	376.0
100.0	264.0	-

Lauric acid ($C_{12}H_{24}O_2$) + Zinc myristate ($ZnC_{28}H_{54}O_4$)

Martin and Pink, 1948

C.S.T. = 114.0°

Lauric acid ($C_{12}H_{24}O_2$) + Zinc stearate ($ZnC_{36}H_{70}O_4$)

Martin and Pink, 1948 (fig.)

%	sat.t.	%	sat.t.
0.1	95	1.0	110
0.2	100	1.5	112
0.5	105	10.0	113

C.S.T. = 113°

Palmitic acid ($C_{16}H_{32}O_2$) + Sodium palmitate ($NaC_{16}H_{31}O_2$)

Donnan and White, 1911

% C	% L	f.t.	
0.65	2.3	60.2	
10.25	4.07	61.1	
13.23	4.96	62.0	
12.57	6.41	63.3	
13.66	7.98	64.4	
15.47	9.91	65.6	
16.32	11.79	65.65	
18.70	13.72	67.75	
20.61	15.56	68.95	
22.84	19.16	70.00	
25.20	22.60	71.00	
29	-	72.3	E (3+1)
34.75	28.59	72.9	
32.21	29.31	73.5	
35.23	28.98	73.5	
35.29	29.98	74.4	
36.08	30.64	76.0	
36.50	32.03	77.6	
36.96	33.36	79.2	
39.64	35.96	82.0	

Ekwall, 1933 (fig.)

mol%	f.t.	E	
0	62.5	-	
2.5	62	60	
10	68	60	
20	72.5	60	
30	76	60	
33.3	82.5	68	(2+1)
37	87.9	72.5	
40	91.5	72.5	
50	123	72.5	
66	-	72.5	
66.7	-	78.5	(1+2)
70	-	85	

Mc Bain and Field, 1933

mol%	f.t.	mol%	f.t.	clear. point
0.00	62.8	38.50	88.4	-
1.75	60.7	39.95	90.0	-
3.64	61.1	41.61	97.8	-
6.65	65.2	43.95	106.2	-
11.19	67.9	46.73	113.2	-
15.55	70.0	50.10	120.5	-
20.74	72.3	53.10	127.8	-
24.04	73.6	57.09	138.0	-
27.47	74.2	61.55	147.8	-
30.34	77.5	66.82	-	168.5
33.35	82.5	71.88	-	203.5
35.99	85.1	100.00	216	316

Vold, Leggett and Mc Bain, 1940 (fig.)

%	clear.p.	%	clear.p.
0	62	70	180
2	60	74	258
10	66	80	278
32	78	85	285
40	88	90	240
43	90	100	297
67	157		

44.5% f.t. = 100°

Leggett, Vold and Mc Bain, 1942

%	f.t.
41.9	90
64.4	150
63.3	200
73.9 (sn)	250
90.2 (sn)	"
91.7 (n)	"
97.2	280

sn = superneat soap .
n = neat soap .

Stearic acid ($C_{18}H_{36}O_2$) + Sodium stearate ($NaC_{18}H_{35}O_2$)

Mc Bain and Stewart, 1933

mol%	clear.p.	mol%	clear.p.
0	68.4	37.5	91.7
2.2	68.6	42.2	107.4
5.0	71.5	42.7	109.6
6.9	72.6	44.4	113.9
12.4	74.7	45.0	118.0
15.6	75.8	50.8	136.0
20.5	77.7	55.4	152.2
23.9	78.1	66.3	188.0
27.0	79.3	66.6	240
30.1	82.0	77.7	242.5
35.0	86.8	100.0	260

E = 67° tr.t.1 = 97° tr.t.2 = 80°

Oleic acid ($C_{18}H_{34}O_2$) + Potassium oleate ($KC_{18}H_{33}O_2$)

Mc Bain and Stewart, 1933

%	mol%	clear.p.	%	mol%	clear.p
0.0	0	13.1	46.55	43.84	84.7
2.500	2.25	13.2	49.16	46.43	97.0
5.491	4.95	12.1	50.62	47.88	103.1
10.87	9.85	11.5	52.35	49.61	120.5
16.13	14.70	18.0	53.79	51.07	134
19.64	17.97	22.4	55.56	52.85	147.7
25.60	23.57	30.4	61.37	58.73	183
29.63	27.39	34.6	66.34	63.84	213
32.43	30.07	37.2	71.63	69.36	244
35.33	32.87	41.7	77.80	75.84	270
39.10	36.53	45.3	81.06	79.36	280
41.61	38.84	56.7	86.10	84.74	287.4
42.60	39.94	63.4	90.36	89.39	287
43.17	40.50	64.0	96.40	96.01	310.5
43.82	41.14	69.0	100.00	100.00	324.4
45.95	43.25	77.6			

Chloracetic acid ($C_2H_3O_2Cl$) + Antimony trichloride ($SbCl_3$)

Usanovich and Sumarokova, 1951

mol%	%	d 50°	d 60°	d 70°
100.00	100.00	2.7352	2.7148	2.6830
90.19	95.64	2.6319	2.6165	2.5918
89.30	95.27	2.6196	2.5973	2.5769
83.10	92.23	2.5519	2.5386	2.5206
77.54	89.29	2.4890	2.4660	2.4479
68.81	84.20	2.3840	2.3621	-
65.21	81.91	2.3419	2.3188	2.2985
56.26	75.55	2.2253	2.2044	2.1803
50.03	70.70	2.1066	2.0815	2.0726
49.08	-	-	-	-
46.48	67.69	2.0987	2.0791	2.0605
45.11	66.51	2.0805	2.0588	2.0420
37.38	59.07	1.9718	1.9554	1.9370
27.85	48.88	1.8309	1.8133	1.7974
14.76	29.48	1.6309	1.6182	1.6000
0.00	0.00	1.3907	1.3777	1.3042

mol%	%	η 50°	η 60°	η 70°
100.00	100.00	3926	3147	2660
90.19	95.64	-	-	2829
89.30	95.27	4500	3615	2898
83.10	92.23	4757	3687	2952
77.54	89.29	4987	3871	3082
68.81	84.20	5160	4075	3211
65.21	81.91	5244	4066	3270
56.26	75.55	5418	4089	3152
50.03	70.70	5457	4109	3169
49.08	-	5449	-	-
46.48	67.69	5272	3971	3155
45.11	66.51	5302	3993	3136
37.38	59.07	5154	3890	2962
27.85	48.88	4894	3690	2865
14.76	29.48	4147	3202	2399
0.00	0.00	3021	2446	2046

Chloracetic acid ($C_2H_3O_2Cl$) + Stannic chloride ($SnCl_4$)

Usanovich, Sumarokova and Glushchenko, 1951

%	mol%	d 50°	d 60°	d 70°
0.00	0.00	1.3907	1.3777	1.3642
5.06	12.81	.4774	.4624	.4499
9.47	22.49	.5513	.5346	.5186
14.91	39.08	.6184	.6014	.5858
18.51	40.97	.6855	.6675	.7775
34.56	60.66	.8228	.8009	.8767
49.28	72.81	.9128	.8980	.8767
66.63	84.62	2.0204	.9998	.9747
77.08	90.33	.0631	2.0418	2.0190
100.00	100.00	.1555	.1347	.1100

mol%	η 50°	η 60°	η 70°
0.00	3091	2446	2051
2.50	3305	2615	2100
5.06	3944	3115	2441
9.51	4193	-	2335
9.57	4100	-	2282
14.91	4170	3013	2214
20.41	3921	2741	1985
25.11	2927	2315	1842
34.56	2003	1796	1450
49.25	1410	1321	1121
66.63	948	874	777
77.08	836	764	680
100	683	650	589

mol%	κ 50°	κ 60°	κ 70°
2.50	1.560	1.767	1.916
5.06	2.621	-	-
7.85	4.092	4.547	4.815
9.51	4.234	4.648	4.877
9.57	-	4.892	-
13.65	4.179	4.761	5.087
14.91	4.125	-	-
18.55	-	4.523	-
20.41	3.549	3.693	-
25.11	3.019	2.693	-
30.46	2.131	1.953	1.735
31.72	1.992	1.956	.872
36.14	1.368	1.439	.530
39.95	0.970	0.830	0.720
49.25	0.290	-	.280
58.52	0.098	-	.090

Trichloracetic acid ($C_2HO_2Cl_3$) + Antimony trichloride ($SbCl_3$)

Sumarokova and Usanovich, 1951

mol%	f.t.	E	tr.t.
100.00	73.2	-	-
80.32	64.0	-	-
71.99	60.0	46.0	-
62.30	56.0	48.0	44.0
59.48	54.5	-	44.0
58.07	53.5	50.5	44.0
49.03	56.0	50.5	-
40.0	54.0	50.5	-
30.07	51.0	-	-
25.40	48.0	40.0	-
21.90	49.5	44.5	-
18.00	50.5	-	-
7.38	54.5	-	-
0.0	58.5	-	-

mol%	%	d 50°	d 60°	d 70°
100.00	100.00	2.7352	2.7148	2.6830
80.32	85.07	2.4784	2.4517	2.4388
58.07	65.92	2.2088	2.1905	2.1755
40.00	48.21	2.0094	1.9915	1.9759
18.00	23.46	1.7870	1.7706	1.7565
0.00	0.00	1.6156	1.6070	1.5864

mol%	η 50°	η 60°	η 70°
100.00	3926	3147	2660
94.61	-	-	-
82.16	-	-	-
80.32	4119	3236	2642
71.99	4179	3312	2687
58.07	-	-	-
49.03	4377	3428	2762
40.00	-	-	-
30.07	4603	3640	2870
18.07	4705	3735	2960
4.21	4806	3811	3113
0.00	4824	3815	3033

mol%	κ 50°	κ 60°	κ 70°
94.61	0.1515	0.1747	-
82.16	0.2005	0.2662	0.3157
80.32	0.2310	0.2474	0.3389
71.89	0.2400	0.2960	-
58.07	0.2150	0.2700	0.3449
49.03	0.1296	0.1648	0.2009
40.00	0.0590	0.0723	0.1077
30.07	0.0268	0.0344	0.0430
18.07	-	-	0.0249

Bromacetic acid ($C_2H_3O_2Br$) + Antimony bromide ($SbBr_3$)

Sumarokova and Khakhlova, 1956

mol%	f.t.	E	mol%	f.t.	E
100	93.0	-	5.52	47.0	-
74.93	86.0	38.5	4.22	47.8	-
49.73	72.0	41.0	1.84	49.3	41.0
35.19	65.0	41.0	0	50.0	-
7.12	46.0	-			

mol%	d 60°	d 70°	d 80°
100	-	-	3.7437
80.08	-	-	3.4520
70.77	-	-	3.3078
59.77	-	-	3.2035
49.64	3.0102	2.9860	2.9781
40.63	2.8029	2.7738	2.7562
27.40	2.5188	2.5023	2.4793
15.26	2.5779	2.2403	2.2217
12.86	2.2056	2.1900	2.1733
2.31	1.9628	1.9333	1.9235
0	1.9103	1.8891	1.8696

mol%	η 60°	η 70°	η 80°
100	-	-	4940
80.08	-	-	4560
70.77	-	-	4380
59.77	-	-	4170
49.64	6060	4710	3830
40.63	5340	4260	3440
27.40	4510	3560	2960
15.26	3790	3110	2580
12.86	3660	3000	2510
2.31	3190	2650	2250
0	3150	2630	2220

Bromacetic acid ($C_2H_3O_2Br$) + Stannic bromide ($SnBr_4$)

Sumarokova and Khakhlova, 1956

mol%	f.t.	E	mol%	f.t.	E
100	32.0	-	35.0	41.8	27.7
90.42	29.5	23.3	23.79	43.7	28.5
87.23	28.8	23.8	21.0	43.8	23.7
74.50	35.0	28.5	9.20	45.6	-
73.48	34.7	28.5	8.17	46.0	-
50.0	40.2	28.0	0	49.8	-
44.40	41.0	28.5			

mol%	η 40°	50°	60°	70°
100	1970	1780	1580	1420
90.42	1880	1650	1470	1320
72.68	2040	1780	1560	1380
47.94	2480	2150	1810	1590
26.36	3070	2580	2150	1820
8.95	3740	3120	2580	2130
0	-	3650	3030	2520

%	d 40°	50°	60°	70°
100	3.3132	3.2894	3.2613	3.2297
96.74	3.2464	3.2199	3.1934	3.1594
89.35	3.0939	3.0701	3.0438	3.0116
74.84	2.8162	2.8005	2.7770	2.7546
53.03	2.4955	2.4812	2.4626	2.4407
23.68	2.1538	2.1394	2.1216	2.1010
0	-	1.9268	1.9103	1.8891

Benzoic acid ($C_7H_6O_2$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911

%	f.t.	E	%	f.t.	E
0	120	-	71.6	60	46
23.0	110	-	78.0	46	-
38.8	100	46	82.0	50	46
50.0	90	46	89.2	60	46
59.0	80	46	97.5	70	46
66.0	70	46	100.0	73	-

Benzoic acid ($C_7H_6O_2$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911

%	f.t.	E	%	f.t.	E
0	120	-	78.4	90	79
20.1	115	-	83.1	85	79
36.8	110	-	87.6	79	-
50.0	105	-	92.0	85	79
61.5	100	79	96.4	90	79
71.0	95	79	100.0	94	-

Phenylacetic acid ($C_8H_8O_2$) + Lithium phenylacetate ($LiC_8H_7O_2$)

Bakunin and Vitale, 1935

mol%	f.t.	E
0.00	76.7	-
9.73	70	-
19.23	70	-
29.43	206	-
42.53	248	59
56.60	271	57
64.52	285	57
76.95	295	58
89.94	306	55.9
98.11	315	58.7
100	320	-

Phenylacetic acid ($C_8H_8O_2$) + Sodium phenylacetate ($NaC_8H_7O_2$)

Bakunin and Vitale, 1935

mol%	f.t.	E
0.00	76.7	-
7.25	70.4	67.2
14.68	70.0	64
22.30	84	62
30.08	91	63
38.08	89	58.8 (2+1)
46.25	78.5	63
54.66	88	62
63.26	111.7	61
67.63	114.7	-
72.11	125	105
81.15	153	102
85.78	158	98
90.47	165	98
95.22	177	95
100.0	186	-

Crawford, 1941

mol %	f.t.		mol %	f.t.
0.0	77.1		64.6	128.9
3.8	74.4		65.4	129.7
6.5	72.2		65.6	129.9
9.0	69.9		66.4	131.1
10.5	68.1		66.6	131.1
11.0	67.9		67.0	131.2
11.1	67.5		67.5	134.5
11.5	68.5		68.0	135.2
11.6	66.9		68.4	137.3
12.6	71.1		69.0	139.0
13.6	73.4		69.6	140.5
14.9	76.3		70.1	142.2
16.4	79.2		70.5	143.5
17.6	81.4		71.0	147.7
18.3	82.6		71.3	144.2
18.9	83.7		71.3	146.2
20.1	85.0		71.7	147.1
21.9	87.5		72.8	150.7
24.0	90.0		73.9	153.3
26.1	91.6		74.6	154.4
27.0	92.3		74.9	155.0
27.4	92.5		75.5	161.4
28.5	93.3		76.5	159.2
30.0	93.9		77.5	161.4
31.5	94.4		78.2	163.0
33.4	94.4	(2+1)	78.6	163.0
35.9	93.9		79.0	165.1
37.4	92.9		79.3	166.0
38.2	91.5		79.7	166.5
39.5	91.4		80.1	166.4
40.9	90.3		80.7	168.0
41.6	90.0		81.5	169.5
42.3	89.4		82.1	169.6
44.0	88.1		83.0	170.8
45.2	86.6		83.6	172.0
45.8	85.8		84.6	174.4
46.5	84.7		85.4	174.7
47.2	83.0		86.1	175.4
47.4	81.6		86.9	177.6
47.7	80.5	E	88.0	179.4
47.9	81.0		89.0	180.1
48.2	81.7		89.5	180.7
48.5	82.9		90.5	182.2
49.0	88.6		91.5	184.5
49.5	92.9		92.2	183.8
50.4	95.9		92.7	186.0
51.6	102.7		93.6	187.2
53.0	108.3		94.0	184.9
54.2	112.2		94.1	187.5
55.5	115.4		94.6	186.3
56.1	117.0		94.8	188.1
57.3	119.3		95.2	187.8
58.0	120.5		96.1	189.6
59.0	122.5		96.6	191.1
59.9	124.2		97.0	191.6
61.0	125.6		98.1	192.9
62.2	126.6		100.0	195.2
62.9	127.4			

mol%	f.t.	m.t.	mol%	f.t.	m.t.
5.9	72	66	46.5	-	80
16.2	76	66	48.0	-	80
26.9	90	65	55.0	-	79
35.5	93	90 (2+1)	59.9	-	75
37.0	-	93	71.0	-	111
38.5	-	91	75.0	-	108
40.0	-	90	79.8	-	105
41.9	-	87	85.0	103	
44.5	-	82	90.0	102	

Phenylacetic acid ($C_8H_8O_2$) + Potassium phenylacetate ($KC_8H_7O_2$)

Bakunin and Vitale, 1935

mol%	f.t.	E	
0.00	76.7	-	
7.78	70	52	
16.33	52	47	
25.09	96	50	
34.25	126	48	
43.86	140	-	(1+1)
53.96	139.5	40	
64.60	135	130	
75.76	161	120	
87.57	189	-	
1000.0	214	-	

Phenylacetic acid ($C_8H_8O_2$) + Rubidium phenylacetate ($RbC_8H_7O_2$)

Bakunin and Vitale, 1935

mol%	f.t.	E	
0.00	76.7	-	
6.40	71	-	
13.35	61	-	
20.91	59	18	
29.15	82	17	
38.14	117.3	-	
48.06	128	16	(1+1)
59.02	121.8	57	
71.17	125	57	
84.76	168	55	
100.0	192	-	

Benzene sulfonic acid ($C_7H_6O_3S$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911-12

%	f.t.	%	f.t.
0	52.5	49.8	5
18.0	45	52.0	13
33.7	35	56.7	25
43.7	25	61.3	35
52.0	13	69.2	45
56.1	5	78.1	55
60.8	-5	90.2	65
47.1	-5	96.4	70
		100	73

Benzene sulfonic acid ($C_7H_6O_3S$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911-12

%	f.t.	E	%	f.t.	E
0	52.5	-	55.2	70	-
15.8	50	44	61.3	75	-
26.2	47.5	44	68.1	80	-
36.9	44	-	77.6	85	-
39.1	50	44	90.3	90	-
45.7	60	-	100	90	-

LVII. METALLIC SALTS + OTHER ORGANIC COMPOUNDS .

Hydrogen (H) + Sodium (Na)

Banus, Mc Sharry and Sullivan, 1955 (fig.)

%	p dissoc. (1+1) (mm^{-2})				
	500°	525°	550°	575°	600°
9.6	0.35	0.75	1.20	2.05	3.95
19.2	.50	1.00	.55	.50	4.10
38.3	"	"	"	"	4.15
57.5	"	"	"	2.55	4.30
76.6	"	"	"	"	4.55
95.8	1.00	1.40	2.05	3.40	-

Hydrogen (H) + Neodymium (Nd)

Mulford and Holley, jr., 1955 (fig.)

Atom%	p			
	600°	650°	700°	800°
86.96	-	-	-	0.55
71.43	0.020	0.050	0.190	1.90
68.97	"	"	"	1.95
50.00	"	"	"	1.95
39.22	"	"	"	1.95
35.71	"	"	"	3.10
35.09	"	"	0.250	-
33.90	"	"	-	-
33.56	0.050	-	-	-

Atom%	p			
	1[a]	2	3	4
32.79	-	-	-	15
32.26	-	-	10	40
31.25	-	5	70	135
30.77	-	30	170	-
30.30	-	60	-	-
29.41	5	250	-	-
28.57	25	-	-	-
27.40	210	-	-	-
	5	6	7	
33.90	-	20	30	
33.33	25	55	165	
32.57	90	385	-	
31.45	415	-	-	

a= the temperatures of isotherms are omitted

Benzene sulfonic acid ($C_7H_6O_3S$) + Antimony trichloride ($SbCl_3$)

Menshutkin, 1911-12

%	f.t.	%	f.t.
0	52.5	49.8	5
18.0	45	52.0	13
33.7	35	56.7	25
43.7	25	61.3	35
52.0	13	69.2	45
56.1	5	78.1	55
60.8	-5	90.2	65
47.1	-5	96.4	70
		100	73

Benzene sulfonic acid ($C_7H_6O_3S$) + Antimony tribromide ($SbBr_3$)

Menshutkin, 1911-12

%	f.t.	E	%	f.t.	E
0	52.5	-	55.2	70	-
15.8	50	44	61.3	75	-
26.2	47.5	44	68.1	80	-
36.9	44	-	77.6	85	-
39.1	50	44	90.3	90	-
45.7	60	-	100	90	-

LVII. METALLIC SALTS + OTHER ORGANIC COMPOUNDS .

Hydrogen (H) + Sodium (Na)

Banus, Mc Sharry and Sullivan, 1955 (fig.)

%	p dissoc. (1+1) (mm⁻²)				
	500°	525°	550°	575°	600°
9.6	0.35	0.75	1.20	2.05	3.95
19.2	.50	1.00	.55	.50	4.10
38.3	"	"	"	"	4.15
57.5	"	"	"	2.55	4.30
76.6	"	"	"	"	4.55
95.8	1.00	1.40	2.05	3.40	-

Hydrogen (H) + Neodymium (Nd)

Mulford and Holley, jr., 1955 (fig.)

Atom%	p			
	600°	650°	700°	800°
86.96	-	-	-	0.55
71.43	0.020	0.050	0.190	1.90
68.97	"	"	"	1.95
50.00	"	"	"	1.95
39.22	"	"	"	1.95
35.71	"	"	"	3.10
35.09	"	"	0.250	-
33.90	"	"	-	-
33.56	0.050	-	-	-

Atom%	p			
	1[a]	2	3	4
32.79	-	-	-	15
32.26	-	-	10	40
31.25	-	5	70	135
30.77	-	30	170	-
30.30	-	60	-	-
29.41	5	250	-	-
28.57	25	-	-	-
27.40	210	-	-	-
	5	6	7	
33.90	-	20	30	
33.33	25	55	165	
32.57	90	385	-	
31.45	415	-	-	

a= the temperatures of isotherms are omitted

Hydrogen (H) + Lanthane (La)

Dreyfus, 1955

Crystallographic study.

Hydrogen (H) + Zirconium (Zr)

Edwards, Levesque and Cubicciotti, 1955

t	atom % (α + δ)	atom % isobaric limit for 760 mm
600	-	34.1
700	51.0 to 40.4	34.9
750	50.4 " 41.1	35.8
800	49.9 " 40.4	36.6
825	49.6 " 40.6	37.2
850	49.3 " 40.7	37.8
850	48.5 to 40.7	37.8
860	48.1 " 40.7	39.8
875	47.8 " 40.7	39.3
885	-	47.4
900	-	48.4

t	p_1			
atom %	60.6	59.7	55.1	52.8
	(β + δ)			
700	4	5.5	7.5	11
750	10.5	15	21	30
800	27	35	51	76
825	42	55	78	120
850	60	83	123	174
860	73	100	145	212
865	80	110	158	242
875	92	128	184	263
900	130	180	252	355

t	p_1			
atom %	45.0	39.3	38.0	36.9
	(β + δ)			
600	-	5	12	54
700	12	36	77	247
750	40	98	245	500
800	127	225	450	695
825	220	378	620	-
850	360	520	737	-
860	450	675	-	-
865	607	-	-	-
875	608	760	-	-

Trzebiatowski and Stalinski, 1956

atom%	lattice constant (Å) α		β	γ	
	400°				
	a	c	a	a	c
100	3.234	5.152	-	-	-
90.91	.233	.150	-	-	-
80.00	.235	.151	4.780	-	-
71.43	.234	.147	.777	-	-
66.67	.232	.149	.781	-	-
63.29	.233	.151	.780	-	-
56.98	.232	-	.779	-	-
51.02	-	-	.779	-	-
46.73	-	-	.780	-	-
43.10	-	-	.779	-	-
40.65	-	-	.779	-	-
39.22	-	-	.778	-	-
38.535	-	-	.780	-	-
37.59	-	-	.781	3.479	4.560
37.04	-	-	.783	.470	.561
36.23	-	-	-	.492	.507
35.778	-	-	-	.496	.500
35.21	-	-	-	.507	.475
34.48	-	-	-	.519	.450
33.44	-	-	-	.519	.450

Hydrogen (H) + Cerium (Ce)

Mulford and Holley, Jr., 1955 (fig.)

atom%	p 100°	200°	300°	400°
31.3	-	-	-	10
30.3	-	-	5	30
29.4	-	-	15	95
28.6	-	5	30	200
27.8	-	10	85	-
27.0	-	25	225	-
26.3	5	80	-	-
26.0	10	230	-	-
25.6	45	-	-	-
25.32	140	-	-	-
	500°	600°	700°	800°
90.9	0.0	-	-	1.0
71.4	0.0	0.23	1.0	5
55.6	0.0	0.23	1.0	6
40.0	0.0	0.23	1.0	6
35.7	0.0	0.23	1.0	8.5
35.09	0.0	0.30	1.0	12
33.90	-	-	20	50
33.3	2	20	55	-
32.3	30	110	-	-
31.3	65	-	-	-
30.3	180	-	-	-
29.85	280	-	-	-

Hydrogen (H) + Plutonium (Pu)

Mulford and Sturdy, 1956 (fig.)

gr Pu/mole H_2	t	p
0.1154	20	250
	100	280
	140	290
	200	300
	210	320
	270	340
	300	370
0.1390	300	250
sic	400	300
	500	340
0.1308	20	150
sic	50	160
	100	170
	140	175
	200	200
	250	230
	300	240
	400	270
	500	300
0.1489	20	50
sic	80	60
	150	70
	180	100
	200	110
	300	140
	400	160
	500	180
0,1641	20	10
	100	15
	200	30
	300	50
	400	110
	500	140
0.1739	180	20
	300	50
	400	80
	500	130
0.1748	100	0
	200	20
	300	30
	400	50
	500	100
0.1847	200	0
	300	20
	400	50
	500	80
0.1993	270	0
	400	20
	450	30
	500	40

Atom %	t	p	atom %	t	p
26.74	150	40	26.13	200	374
26.77	188	110	25.64	150	516
26.88	207	215	27.32	330	710
26.95	207	215	27.03	320	325
26.91	223	330			

Mulford and Sturdy, 1955 (fig.)

t	p dissoc. (2+1)	t	p dissoc. (2+1)
400	0.006	700	40
500	0.28	800	260
600	4.5		

Deuterium (D) + Plutonium (Pu)

Mulford and Sturdy, 1956 (fig.)

Atom%	p	Atom%	p	Atom%	p
100°		200°		300°	
26.60	30	28.73	0	28.99	30
25.64	140	27.47	70	28.33	160
25.51	230	27.32	120	27.93	340
		27.03	260	27.78	450
		26.67	320		
400°		500°			
31.25	0	33.3	0		
29.85	100	30.58	170		
29.24	250	30.12	320		
29.07	280	29.85	490		
28.90	430	29.59	580		
28.57	520				

Atom%	p	Atom%	p
200°			
28.57	0	74	400
28.25	10	67	450-400
27.78	40	53	520-500
27.40	90	46	520
27.25	100	32	670-650-620
27.03	180	25	650-590-540
26.95	150	18	740
26.88	270	03	790
26.81	290		

Mulford and Sturdy, 1955 (fig.)

t	p dissoc.
(2+1)	
600	5.8
700	50
800	300

Hydrogen (H) + Praseodymium (Pr)

Mulford and Holley, Jr.,1955 (fig.)

Atom%	p 200°	300°	400°	500°
32.3	-	-	-	20
31.3	-	-	10	60
30.3	-	5	45	130
29.4	5	20	100	220
28.6	20	65	200	350
27.7	35	150	360	-
27.0	75	275	-	-
26.3	200	-	-	-
25.6	400	-	-	-

Atom%	600°	650°	700°	750°	800°
94.94	0.010	0.030	0.050	0.15	0.20
90.91	.015	.045	.100	.22	.40
83.33	"	.080	.20	.50	1.00
76.92	"	"	.30	.90	2.5
71.43	"	"	"	1.10	3.0
50.00	"	"	"	"	"
37.04	"	"	"	"	4.0
34.48	"	"	0.50	"	-
32.26	50	-	90	-	-
31.25	120	-	185	-	-
30.30	210	-	300	-	-
29.41	340	-	360 (2.35)	-	-

400°		500°	
33.33	5	33.56	10
32.79	60	32.89	150
32.26	170	32.79	240
31.85	250	32.47	280

600°		650°	
78.74	0.011	83.33	0.01
54.05	.011	75.76	0.05
35.46	.011	35.71	0.05
33.90	1.500	34.48	4.00

700°		750°	
74.07	0.1	84.75	0.08
64.52	0.12	71.43	0.5
47.62	0.12	62.50	0.8
35.46	0.12	40.00	0.8
34.48	4.00	35.71	0.9
		34.48	8

800°	
92.59	0.11
71.43	1.0
44.44	2.0
37.74	2.0
35.09	9.0

Hydrogen (H) + Gadolinium (Gd)

Sturdy and Mulford,1956 (fig.)

Atom%	p	Atom%	p
200°		225°	
29.07	20	32.78	0
26.53	20	31.47	20
25.77	40	29.41	40
25.64	80	27.03	40
25.58	240	25.91	50
		25.64	140
		25.58	250
250°		300°	
32.78	0	32.78	0
30.49	60	31.25	80
27.70	60	30.49	180
26.11	60	28.99	180
25.64	240	27.03	180
		25.77	220
		25.71	250

Sodium fluoride (NaF) + Sodium (Na)

Bredig, Johnson and Smith, jr., 1955

mol%	f.t.	sat.t.	mol%	f.t.	sat.t.
0	995	-	80.5	990	1026
0.19	979	-	82.7	"	1004
0.96	983	-	83.6	983	-
7.6	990	1000	85.4	976	-
14.8	"	1015	89.3	952	-
15.8	"	1001	91.9	900	-
76.4	"	1078	94.7	849	-
80.0	"	1048	97.6	801	-

Potassium fluoride (KF) + Potassium (K)

Bredig, Bronstein and Smith, jr., 1955

mol% L_1	mol% L_2	sat.t.	mol% L_1	mol% L_2	sat.t.
-	51.1	853	-	56.9	865
6.0	-	862	14.3	-	877
5.8	50.4	"	14.8	-	"
6.3	-	863	15.6	49.2	881
7.4	55.0	865	18.0	50.5	886
9.7	58.0	"	-	48.1	"
-	51.0	"	21.7	44.0	894
-	50.7	"	-	42.4	"
-	53.0	"	30.5	46.0	905

Cesium fluoride (CsF) + Cesium (Cs)

Bredig, Bronstein and Smith,jr., 1955

mol%	f.t.	mol%	f.t.
3.9	349	63.3	662
11.4	452	74.6	671
22.1	527	93	692
36.3	599	100	703
50.1	640		

Sodium chloride (NaCl) + Sodium (Na)

Bredig, Johnston and Smith jr., 1955

mol %	f.t.	sat.t.
0	803	-
0.08	750	-
0.11	784	-
0.15	790	-
2.8	"	809
4.9	"	826
9.3	"	889
10.2	"	890
21.3	"	929
25.6	"	949
25.1	"	950
27.8	"	975
32.0	"	976
33.0	"	1001
88.9	"	957
91.7	"	913
94.2	"	874
95.8	"	832
97.1	"	-
97.6	784	-
97.8	781	-
98.2	750	-
99.25	649	-
99.88	550	-

Chlorine (Cl_2) + Potassium iodide (KI)

Ephraim, 1917

t	p dissoc. (1+1)	t	p dissoc.
41	28	102	292
58	58	112	387
71.5	114	122	505
79.5	153	131	620
90	210	136	704

Chlorine (Cl_2) + Rubidium bromide (RbBr)

Ephraim, 1917

t	p dissoc. (1+1)	t	p dissoc.
17	20	83.5	455
57	75	91.5	699
72.5	235	93	760

Chlorine (Cl_2) + Rubidium iodide (RbI)

Ephraim, 1917

t	p dissoc. (1+1)	t	p dissoc.
97	80	141.5	442
109	146	146.3	494
120	235	155	555
134	403	161	600

Cesium chloride (CsCl) + Cesium (Cs)

Bredig, Bronstein and Smith jr., 1955

mol%	f.t.	mol%	f.t.
94.1	486	43.9	611
91.4	513	22.0	621
84.5	549	0	640
67.6	601		

Chlorine (Cl_2) + Cesium bromide (CsBr)

Ephraim, 1917

t	p dissoc.	t	p dissoc.
	(1+1)		
79	45	132.5	612
99	125	135.3	680
112	228	138	760
124.5	420		

Chlorine (Cl_2) + Cesium iodide (CsI)

Ephraim, 1917

t	p dissoc.	t	p dissoc.
	(1+1)		
112	24	193	455
134	70	199.5	555
147	103	204.5	640
157	140	206	680
166	185	207	705
169	203	209	760
178.5	273		

Cadmium chloride ($CdCl_2$) + Cadmium (Cd)

Urazov and Karnaukhov, 1954

wt%	mol%	f.t.		E
		$CdCl_2$	Cd	
0	0	568	-	-
1.49	2.49	-	-	564
2.17	3.45	-	-	556
4.99	7.81	543	-	520
8.05	12.54	530	-	"
9.08	14.04	-	-	"
9.9	15.06	-	-	"
10.59	16.04	-	326	"
15.0	22.17	-	"	"
20.0	28.74	550	"	"
25.0	34.96	549	"	"
30.0	40.88	"	325	521
40.0	51.81	"	326	519
50.0	61.73	"	"	518
55.0	66.48	550	"	"
60.0	72.22	549	"	"
65.0	74.97	547	"	"
70.0	79.03	549	"	521
85.0	90.15	"	"	"
90.0	93.64	"	"	"
95.0	96.83	-	"	520
100.0	100.00	-	"	-

Lead chloride ($PbCl_2$) + Lead (Pb)

Urazov and Karnaukhov, 1954 and 1956

%	f.t.		%	f.t.	
	Pb	$PbCl_2$		Pb	$PbCl_2$
0	-	497	50	327	486
5	327	489	60	"	"
10	"	486	75	"	"
20	"	"	80	"	"
25	"	"	100	"	-
40	"	"			

Chlorine (Cl_2) + Potassium chloraurite ($KAuCl_2$)

Parravano and Malquori, 1926

t	p dissoc.	t	p dissoc.
	(1+1)		
250	8.90	302	39.80
256	9.12	322	67.60
267	16.70	315	16.62
381	23.00	390	91.50
300	39.10		

Chlorine (Cl_2) + Cesium chloraurite ($CsAuCl_2$)

Parravano and Malquori, 1926

t	p dissoc.	t	p dissoc.
	(1+1)		
302	12.60	360	45.88
309	15.70	390	91.50
315	16.62	400	114.00
350	37.90		

Chlorine (Cl_2) + Silver chloraurite ($AgAuCl_2$)

Parravano and Malquori, 1926

t	p dissoc.	t	p dissoc.
	(1+1)		
100	4.17	162	22.23
110	5.47	180	36.00
140	12.34	194	63.10
150	16.10	213	85.50

Chlorine (Cl) + Bismuth (Bi)

Herz and Guttmann, 1907

mol%(3+1)	f.t.	mol%(3+1)	f.t.
	193°		
94.3	177	66.3	158
84.3	164	64.5	225
79.4	156	54.9	267
77.5	162	100	227
68.6	163		

Eggink, 1908

atom%	f.t.		m.t.	
	cooling	heating	cooling	heating
100	271.5	271.5	271.5	271.5
95	-	-	267.5	267.5
94.30	269	-	-	-
93.60	283	-	-	-
89.52	303	-	-	-
85	-	-	267.5	267.5
83.82	332	-	-	-
82.5	340	-	-	-
75	315	321	267.5	257.5
65	316	321	266.5	287.5
55	316	320	262.5	267.5
50	317	320	262.5	267.5
48	316	321	-	-
45	316	320	-	-
42	315.5	320	209	209.5
38	315	321	208	209
35	-	-	208	209
33.9	328	-	-	-
33.3	-	-	208.8	209
31.8	334	-	-	-
30.3	291	-	-	-
30.2	264	-	-	-
30	-	-	208	209
28.9	236	-	-	-
28.1	210	-	-	-
26.5	224	225	208	209
25	232.5	232.5	232.5	232.5
22.4	-	220	-	204
20.8	-	221	-	206
20	-	226	-	208
18.4	-	218	-	-
16.2	-	213	-	-
14.2	-	206	-	-

Marino and Becarelli, 1916

atom% Bi	wt% $BiCl_3$	f.t.	E
25	100	272	-
31.9	95	264	-
37.8	90	"	-
47.4	80	261	190
47.8	70	260	178
61.1	60	256	182
66.0	50	"	190
70.0	40	255	"
73.5	30	249	"
76.4	20	237	197
79.0	10	212	-
96.6	0	224	-

Graff, 1933

%(3+1)	f.t.	%(3+1)	f.t.
0	-103	60	-132
10	-106	60	-135.4
20	-108.5	70	-133
30	-113	80	-124
40	-118	100	-108.7
50	-124		

Aten, 1909

mol% (3+1)	d		
	260°	270°	280°
100	3.85	3.84	3.83
90.7	4.06	4.04	4.02
81.4	-	-	4.32

mol% (3+1)	d		
	290°	300°	310°
100	3.81	3.79	3.77
90.7	4.00	3.98	3.94
81.4	4.25	4.20	4.18
69.8	-	4.59	4.54

mol% (3+1)	d		
	320°	330°	340°
100	3.34	3.71	3.67
90.7	3.91	3.89	3.87
81.4	4.16	4.14	4.13
69.8	4.52	4.49	4.47
59.8	4.89	4.86	-

Aten, 1909

mol% (3+1)	η 260°	η 270°	η 280°
100	32000	29500	27000
90.7	41000	37000	33500
81.4	-	-	47000

mol% (3+1)	η 290°	η 300°	η 310°
100	25000	23000	21500
90.7	30500	28500	26000
81.4	42500	38000	34500
69.8	-	61000	54500

mol% (3+1)	η 320°	η 330°	η 340°
100	20500	19000	18000
90.7	24500	23500	22000
81.4	32000	30000	28500
69.8	49500	45000	41500
59.8	87500	77500	-

mol% (3+1)	κ 250°	κ 260°	κ 270°
100	4060	4280	4470
94.8	4010	4230	4440
89.9	3850	4110	4360
84.8	3600	3910	4210

mol% (3+1)	κ 280°	κ 290°	κ 300°
100	4640	4820	4980
94.8	4660	4870	5060
89.9	4610	4850	5070
84.8	4500	4760	5010
79.5	4260	4550	4810
74.1	-	4210	4480

mol% (3+1)	κ 310°	κ 320°	κ 330°
100	5120	5240	5350
94.8	5250	5410	5350
89.9	5280	5500	5700
84.8	5250	5490	5720
79.5	5070	5330	5590
74.1	4760	5030	5300
68.9	4350	4650	4940
63.4	3990	4310	4620
57.9	(3710)	3910	4200
51.7	-	(3870)	(4110)

mol% (3+1)	κ 340°	κ 350°
100	5450	5540
94.8	5710	5850
89.9	5890	6050
84.8	5950	6170
79.5	5830	6070
74.1	5570	5850
68.9	5270	5490
63.4	4900	5130

mol% (3+1)	κ (in mhos) 290°	κ (in mhos) 300°	κ (in mhos) 310°
0	7760	7720	7690
0.23	7640	7620	7590
0.58	7350	7330	7300
0.81	7130	7100	7070

mol% (3+1)	κ (in mhos) 320°	κ (in mhos) 330°	κ (in mhos) 340°	κ (in mhos) 350°
0	7650	7620	7580	7550
0.23	7560	7520	7490	7460
0.58	7270	7250	7220	7190
0.81	7040	7020	6990	6970
1.08	-	6670	6640	6610

Chlorine (Cl_2) + Chromium oxychloride (CrO_2Cl_2)

Bakhuis Roozeboom, 1885

mol%	p	mol%	p
	0°		
54.0	1302	63.3	1069
58.3	1208	64.1	1051
58.8	1198	65.8	1016
59.7	1189	66.7	984
60.6	1150	68.7	923
61.5	1125	71.9	824
62.5	1101	74.6	753

mol %	f.t.	mol %	f.t.
74.1	0	46.5	-21
61.7	-14	40.0	-24

t	p 39.4 mol%	42.0 mol%	46.5 mol%	50.0 mol%
-10	-	-	-	1015
-12	-	1105	-	-
-14	-	-	-	902
-15	-	1002	-	-
-16	-	-	-	851
-18	-	905	831	-
-20	896	862	780	-
-22	844	820	734	-
-24	791	-	-	-
-26	765	-	-	-

t	p 54mol%	58.8mol%	61.7mol%	65mol%	72.5mol%
0	1295	1185	1118	1031	825
-2	1232	-	-	-	-
-3	-	-	1034	-	-
-5	-	-	975	-	-
-7	-	948	921	-	-
-8	-	930	-	-	-
-9	995	-	870	-	-
-10	-	-	-	-	-
-11	952	-	818	-	-
-12	-	-	-	-	-
-14	-	-	754	-	-

Biltz and Meinecke, 1923

mol %	f.t.	E
100	-22.5	-
93.2	-26.5	-110
87.0	-30.5	-108
78.5	-36.0	-108
67.2	-44	-107.5
57.7	-51	-108
57.0	-53	-108
50.3	-56	-
47.6	-58	-108
37.2	-71.5	-107.5
35.9	-71	-108
31.6	-79.5	-108
25.0	-89.5	-108
22.8	-96	-108.5
22.3	-96	-108
8.5	-105	-107.5

Kordes, 1926

mol %	f.t.
0	- 23
13	-108 E
100	-102

Chlorine (Cl_2) + Titanium tetrachloride ($TiCl_4$)

Krieve and Masou, 1956

mol%	p_1	mol%	p_1
20°		30°	
74.87	1.50	57.73	3.39
64.15	2.50	51.56	3.93
56.42	2.71	44.29	4.63
50.36	3.13	35.88	5.41
43.26	3.67	28.93	6.14
35.01	4.24		
27.71	4.76		

Chlorine (Cl_2) + Tin tetrachloride ($SnCl_4$)

Biltz and Meinecke, 1923

mol%	f.t.	E
100	- 33.0	-
96.5	- 34.0	-
92.6	- 36.0	-
82.8	- 40.0	-
81.3	- 42.5	-
75.6	- 46.5	- 104
70.2	- 50.0	-
67.6	- 52.2	- 105
63.4	- 56.0	- 105
57.3	- 61.0	- 105
52.5	- 63.5	- 104
52.3	- 66.0	- 105.0
47.3	- 71.5	- 104.5
44.2	- 75.5	- 105.0
39.7	- 79	- 105.0
37.0	- 82	-
34.7	- 89	- 105.0
31.8	- 90	- 104.5
31.2	- 96	- 106.5
28.9	- 99	- 105.0
23.8	-	- 106.0
21.9	-	- 106.0
19.0	-	- 106

Chlorine (Cl_2) + Platinum (Pt)

Shchukarev, Oranskaya and Shemyakina, 1956

t	p dissoc.	t	p dissoc.
	(4+1)		
298	7.5	327	34.5
306	13.5	332	44.5
308	14.5	332	36
312	19.0	336	59.5
312	20.5	337	60
316	21.0	338	54.5
318	21.5	342	65
318	23.5	342	69
318	22.0	346	83.5
322	31.0	348	109
322	31.5	353	119
322	30.5	358	150
(3+1)		(2+1)	
332	7	490	34
348	14	-	36
"	"	494	32
"	"	496	43
"	15	500	45
350	17	"	49.5
352	20	506	59
364	28	"	64.5
"	26	510	68
372	57.5	515	77.5
394	77	520	91
		526	130.5
		530	144.5
		"	154.5
	(1+1)		
568	14.5	734	152
580	25	752	172
592	24	762	201
614	23	762	186
700	83		

Sodium bromide (NaBr) + Sodium (Na)

Bredig, Johnson and Smith,jr., 1955

mol%	f.t.	sat.t.	mol%	f.t.	sat.t
0	747	-	14.6	740	953
0.04	650	-	81.9	"	955
0.07	685	-	92.2	"	894
0.15	720	-	95.4	"	825
4.6	740	750	96.7	"	757
6.8	"	799	97.6	721	-
6.4	"	849	98.7	687	-
7.3	"	852	98.9	660	-
12.2	"	909	99.59	600	-
14.7	"	943	99.79	550	-

Bromine (Br_2) + Potassium iodide (KI)

Ephraim, 1917

t	p dissoc.	t	p dissoc.
	(1+1)		
17	5	71	98
40	15	80	140
54.5	47		

Bromine (Br_2) + Rubidium chloride (RbCl)

Ephraim, 1917

t	p dissoc.	t	p dissoc.
	(1+1)		
18	30	77	560
50	80	80	700
62	240	81	760
69	340		

Bromine (Br_2) + Rubidium iodide (RbI)

Ephraim, 1917

t	p dissoc.	t	p dissoc.
	(1+1)		
106	45	165	385
125.5	111	174.5	505
137	162	180	600
147	230	186.5	760
156	300		

Bromine (Br_2) + Cesium chloride (CsCl)

Ephraim , 1917

t	p dissoc.	t	p dissoc.
	(1+1)		
54	15	114	482
77	90	120	624
91	170	122.5	685
104	317	124	760

Bromine (Br_2) + Aluminum bromide etherate ($AlBr_3.C_4H_{10}O$)

Plotnikov and Kaplan, 1948

mol%	κ	mol%	κ
2.60	0.10	16.00	9.64
5.80	1.66	16.00	9.69
7.00	3.91	17.30	10.60
9.02	4.68	22.50	12.10
10.2	5.28	23.90	11.70
11.10	6.45	25.00	12.50
12.20	7.11	29.00	13.20
12.25	7.11	34.70	13.00
12.90	7.70	36.70	13.50
13.80	8.21	39.70	13.10
14.90	8.65	43.00	12.40
15.20	9.80	45.30	11.40

Bromine (Br_2) + Cesium iodide (CsI)

Ephraim, 1917

t	p dissoc.	t	p dissoc.
	(1+1)		
157	45	219	342
170	85	231	506
189.5	160	235	600
206	240	242.5	760

Bromine (Br_2) + Aluminum bromide ($AlBr_3$)

Pushin and Makucz, 1938

mol %	f.t.	E
100	97	-
82.5	88	-17
74	82.5	-15
56.5	67.5	-14.5
44.5	53	-14.5
33.5	36	-14
25	22	-13.5
20.5	10	"
14	- 9	"
8.8	-13.5	"
6.5	-10.5	"
0	- 7.3	-

Bromine (Br_2) + Aluminum chloride etherate ($AlCl_3.C_4H_{10}O$)

Plotnikov and Kaplan, 1948

mol %	κ	mol %	κ
6.1	2.7	30.0	17.6
12.2	8.8	34.6	18.4
16.3	12.2	37.2	19.6
18.0	13.2	41.7	18.6
24.2	17.0	46.6	19.6
29.3	18.0	65.0	16.4

Bromine (Br_2) + Complex ($AlBr_5.C_2H_5Br.CS_2$)

Plotnikov, 1903

% (complex)	κ	% (complex)	κ
	18°		
0.27	very low	23.7	54
0.56	13.10^{-3}	24.4	54
1.4	81.10^{-3}	25.0	60
17.8	40	26.3	56
18.1	50	29.2	63
21	54	29.6	64
21.5	53	31.0	64
23.6	54	33.1	64

M	λ	M	λ
	18°		
1.62	4.1	1.20	4.5
1.52	4.3	1.16	4.7
1.45	4.4	1.05	5.0
1.43	4.4	1.05	5.2
1.32	4.3	0.90	5.6

Bromine (Br_2) + Complex ($AlBr_7 . CS_2$)

Plotnikov, 1903

% (complex)	ϰ	% (complex)	ϰ
	18°		
3.0	38.10^{-4}	15.7	57
3.24	40.10^{-4}	15.9	55
3.88	"	17.0	57
4.2	38.10^{-4}	21.1	"
4.4	32.10^{-4}	23.7	"
4.5	30.10^{-4}	24.7	"
4.59	40.10^{-4}	29.2	"
5.4	34.10^{-4}	33.6	55
6.0	"	21.5	57
8.2	32.10^{-4}	23.3	"
8.7	34.10^{-4}	25.0	"
9.9	33.10^{-4}	27.9	58
10.3	27	31.9	"
11.5	37	33.6	62
12.6	"	35.0	"
14.0	51	43.0	"
14.2	53	45.7	61
14.5	55	saturated	64
15.5	58		

Bromine (Br) + Platinum (Pt)

Shchukarev, Tolmachev and al., 1956

t	p dissoc.	t	p dissoc.
	(4+1)		
200	12	250	98
210	19	260	143
220	29	270	205.5
230	44	280	291
240	66.5		
	(3+1)		
280	24	340	177
290	34	350	228
300	48	360	302
310	67	370	396
320	92	380	515
330	126	390	665
	(2+1)		
420	36	470	112
430	46	480	139
440	58	490	170
450	72	500	20
460	90		
	(1+1)		
460	18	490	32
470	22	500	38
480	26.5	510	46

Bromine (Br) + Bismuth (Bi)

Herz and Guttmann, 1907

atom %	f.t.	atom %	f.t.
	(3+1)		
25.0	215	49.7	198
44.7	194	50.3	196
48.3	191	56.7	188

Eggink, 1908

atom%	m.t.	f.t.	E
100	271.5	271.5	-
97	266.5	266.5	263.5
95	-	-	262.5
90	-	-	262.5
85	-	-	262.5
80	281.5	281.5	261.5
75	283	286	260
65	283.5	286	261.5
55	283.5	286.5	261.5
50	284	286	263.5
48	285	287	263.5
45	284	286	-
42	284	286	205
38	284	287	205
35	-	-	204
33	-	-	204
30	-	-	204
27	214	215	204
25	217.5	218.5	-

%	sat.t.	%	sat.t.
41.1	306	84.5	299
39.8	297	87.7	311

Marino and Becarelli, 1913

%	f.t.	E
100	272	-
95	256	-
90	253	-
80	248	200
75	247	200
70	246	202
65	237	202
60	235	203
55	200	160
50	210	155
46.2	210	153
41.4	192	152
39.5	181	152

Aten, 1905

atom %	f.t.	m.t.
100.0	277	277
99.1	318	271
95.8	452	272
82.8	535	263
71.2	602	267
59.3	656	270
53.1	702	-
47.6	760	-

Bromine (Br_2) + Uranium hexafluoride (UF_6)

Fischer and Vogel, 1954

mol%	f.t.	m.t.	E by cooling	E by heating
0	-7.3	-	-7.3	-
1.23	0	-	-7.4	-7.4
2.21	10.3	-	-7.5	-7.5
3.08	19.9	-	-7.4	-7.4
4.84	31.7	31.3	-7.3	-7.3
7.68	38.8	39.9	-7.4	-7.4
13.16	45.1	46.2	-7.5	-7.5
19.17	46.9	47.6	-7.4	-
34.53	48.5	48.8	-7.5	-7.2
48.96	48.7	-	-7.2	-7.2
54.71	50.0	-	-	-
61.70	50.9	50.4	-7.5	-7.4
71.78	52.8	-	-	-
74.80	53.2	52.4	-7.3	-7.3
79.70	54.6	-	-	-
84.16	56.1	-	-7.2	-7.0
85.01	56.9	-	-	-
89.94	58.3	-	-	-
93.66	60.6	60.9	-	-
100	63.9	-	-	-

Sodium iodide (NaI) + Sodium (Na)

Bredig, Johnson and Smith Jr., 1955

atom %	f.t.	sat.t.
0	661	-
4.1	656	722
6.2	"	763
8.3	"	810
9.9	"	849
14.9	"	909
20.8	"	954
89.5	"	959
94.3	"	882
97.2	"	800
98.5	"	699
98.8	652	-
99.20	637	-
99.40	608	-
99.72	584	-
99.94	550	-

Ceisum iodide (CsI) + Cesium (Cs)

Bredig, Bronstein and Smith Jr., 1955

mol %	f.t.
98.3	399
95.7	450
81.0	549
63.4	571
44.9	580
27.5	592
0	626

Iodine (I_2) + Cesium bromide (CsBr)

Ephraim, 1917 .

t	p dissoc.
(1+1)	
172	290
182.5	395
191	511
200	655

Iodine (I_2) + Tetramminezinc iodide ($ZnN_4H_{12}I_2$)

Fialkov and Shevchenko, 1952

mol %	f.t.	E
0.54	111.8	-
1.28	111.2	-
3.92	108.8	-
5.03	107.1	55.8
5.57	106.2	55.9
6.49	103.4	55.9
7.57	98.7	56.0
9.62	88.3	56.1
10.29	-	"
11.36	-	"
11.81	-	56.2
12.74	-	"
14.74	-	"
15.93	-	"
16.43	-	56.1
17.66	79.8	56.0
18.54	83.0	-
19.65	90.8	58.3
20.12	93.1	58.2
20.84	93.9	58.3
22.00	104.6	"
24.36	115.3	58.1
25.40	119.8	58.2
27.20	125.2	58.7

Fialkov and Shevchenko, 1952

mol%	d	mol%	d
	130°		
0	3.9192	6.50	3.8238
1.91	.9037	10.56	.7276
3.57	.8600	14.89	.6602
4.82	.8493		

mol %	κ 130°	κ 140°
0	0.207	0.196
0.31	14.6	14.2
0.71	82.8	82.2
1.17	151	-
2.03	375	370
3.32	766	768
4.67	1150	1200
7.40	1480	1510
10.46	1850	1870
12.92	1880	1960
15.59	1860	-
17.54	1870	1910
19.37	1870	-

Iodine (I_2) + Tetramminecadmium iodide ($CdN_4H_{12}I_2$)

Fialkov and Shevchenko, 1952

mol.%	f.t.	E
0.48	111.2	-
1.03	110.3	-
1.94	108.4	-
2.76	106.9	-
3.61	105.0	-
5.17	99.7	68.9
7.83	90.1	70.2
9.47	-	70.2
10.90	-	70.3
12.26	-	70.4
14.10	-	70.3
15.23	-	72.4
16.02	87.8	-
16.95	89.3	74.6
17.91	-	73.9
18.44	90.6	74.4
19.71	92.2	74.1
20.92	97.3	74.4
22.30	101.5	-
23.12	107.9	-
24.86	124.6	-

mol%	d	mol%	d
	130°		
0	3.9181	11.28	3.7757
2.76	.8870	15.51	.7069
5.82	.8506	25.55	.5538

mol %	κ 130°	κ 140°
0.19	3.49	3.26
0.73	31.6	29.8
1.57	136	132
2.44	225	214
3.20	439	441
4.34	621	636
6.94	926	952
9.12	1150	1220
11.17	1310	1370
13.76	1420	1480
15.80	1460	1560
18.48	1470	1570
22.22	1450	-
24.94	1460	1550

mol %	λ	mol %	λ
0.19	11.83	9.12	92.25
0.73	28.42	11.17	88.36
1.57	57.21	13.76	80.98
2.44	61.64	15.80	74.82
3.20	92.47	18.48	67.15
4.34	98.03	22.22	58.31
6.94	94.66	24.94	54.51

Iodine (I_2) + Tetramminenickel iodide ($NiN_6H_{18}I_2$)

Fialkov and Shevchenko, 1950

mol%	f.t.	E	
0.0	112.5	112.5	
4.5	104.1	79.4	
6.0	99.4	85.8	
7.0	95.3	86.4	
8.0	89.2	"	
8.5	-	"	
8.9	-	85.5	
9.0	-	86.4	E
9.5	93.3	86.3	
10.0	96.5	86.3	
10.5	97.0	-	
11.0	97.5	-	(8+1)
11.2	95.8	-	
11.5	96.6	-	
12.8	-	94.3	E
13.0	-	94.2	
13.9	117.0	94.3	
15.1	-	93.2	
16.0	131.9	94.3	
17.0	133.5	-	
18.0	135.3	93.0	
18.8	135.4	93.5	
21.3	133.4	-	

mol%	d
	130°
0	3.9103
1.07	.8811
3.93	.7994
5.93	.7393
10.42	.5897

mol %	κ 130°	140°
0	0.178	0.140
0.22	10.4	10.8
0.56	51.9	52.7
0.81	96.0	96.2
1.28	238	247
2.03	522	544
2.17	551	561
3.02	973	1040
4.30	1498	1550
6.27	1860	2030
8.23	1990	2020
9.08	2130	2160
12.62	2130	2180
12.81	2120	-

Iodine (I_2) + Bismuth (Bi)

Marino and Becarelli, 1913

%	f.t.	E	tr.t.	E_3
100	285	-	-	-
95	339	-	294	284
90	337	337	294	284
80	340	340	294	282
70	340	340	293	280
66	342	342	294	280
62	338	338	294	282
58.5	344	344	295	284
55	344	344	294	283
50	352	344	294	283
45	374	344	296	284
43	379	342	294	284
40	394	340	296	284
35.3	412	-	-	-
30	404	117	-	-
20	-	117	-	-
10	292	117	-	-

Van Klooster, 1913

%	f.t.	E	tr.t.	E_3
100	272	-	-	-
99	270	-	-	-
96.9	-	321	277	268
83	-	334.5	281	266
71.2	-	339	280	268
62.1	-	338.5	281	269
52.2	-	338	278	268
45	367	338	281	268
41.2	380	335	278	268
38.6	393	334	273	-
35.3	408	-	-	-
17.7	122	113	-	-
1.6	113	-	-	-
0	113.5	-	-	-

(3+1) (1+1)

Iodine (I_2) + Aluminum iodide (AlI_3)

Nijnik, 1937

wt%	mol%	f.t.	E
100	100	188.0	-
80.78	72.34	150.1	93.5
68.65	57.70	131.9	93.6
53.35	41.60	111.5	93.6
55.74	43.92	109.4	93.5
52.17	40.44	106.6	93.8
45.70	35.40	95.6	-
39.15	28.62	95.7	93.5
30.08	21.15	100.5	93.0
21.44	14.54	102.4	93.6
13.44	8.85	104.7	-
4.02	2.59	109.7	93.7
0	0	112.7	-

Iodine (I_2) + Indium(IN)

Peretti, 1956 (fig.)

atom%	f.t.	E		
100	156.4	-		
99.8	365	155		
84	"	"		
70	"	"		
60	"	"		
50	"	216	(1+1)	$L_1 + L_2$
45	350	"		
40	330	"		
33.4	224.6	"	(2+1)	
31.8	216	"		
31	220	147		
28.2	147	"		
25	207	96.8	(3+1)	
20	150	"		
15	125	"		
9.8	96.8	"		
5	105	-		
0	113.6	-		

Iodine (I) + Gallium (Ga)

Corbett and Mc Mullan, 1955 (fig.)

atom%	f.t.	atom%	f.t.
50	265	33	210
45	265	30	180
40	255	29.5 E	160
35	220	28	200
34. E	205	25	215

(1+1) (1+2) (1+3)

Iodine (I) + Tin (Sn)

Reinders and Lange, 1913

% (SnI_4)	b.t.	% (SnI_4)	b.t.
0	183	70	214
10	184	75	219.5
20	187	80	228
30	190	85	240
40	193	90	267
50	198	95	296
60	204	100	340

% (SnI_4)	f.t.	% (SnI_4)	f.t.
0	113.2	60	79.6 E
10	109.0	65	83.5
20	104.7	70	89.8
30	99.7	80	108.4
40	94.7	90	127.0
50	87.6	100	143.5
55	83.0		

Plotnikov, Fialkov and Chalii, 1936

%	mol %	$\varkappa$ 130°	$\varkappa$ 140°
0	0	0.319	0.303
1.26	0.51	0.297	0.288
2.82	1.16	0.276	0.267
10.95	4.75	0.177	0.172
18.97	8.67	0.107	0.104

(see also SnI_4) .

Van Klooster, 1910

%	f.t.	tr.t.
0	113.5	-
1.3	111.5	-
8.8	90	-
10.7	79	-
14.5	102.5	-
17.6	133	-
19	143	-
21.25	216	-
21.25	321	143.5
25.4	317	-
31.9	314	-
31..9	321	-
31.9	321	-
48.4	321	-
73.8	317.5	-
89.3	310	-
95.85	?	-
100.00	232	-

Oxygen (O) + Sodium (Na)

Rode and Golder, 1956

atom%	f.t.	atom%	f.t.
33.3	100	49.3	284
33.6	122	50.0	294
33.7	143	"	304
33.9	152	"	354
34.0	164	50.3	386
34.2	190	50.5	408
34.5	208	"	432
34.7	221	50.8	452
35.1	230	51.5	488
35.3	244	52.1	510
35.7	254	52.9	546
37.6	262	53.8	554
38.6	266	54.3	558
43.1	272	55.9	568
45.0	274		
47.4	280		

Oxygen (O) + Cadmium oxide (CdO)

Gilbert and Kitchener, 1956

Dissociation of CdO

p_2	t	p_2	t
0.132	880	0.939	968
.179	890	1.42	1001
.218	903	1.65	1006
.345	927	1.72	1009
.328	928	1.69	1010
.375	934	3.33	1051
.630	958	3.51	1053
.740	963	3.51	1053
.865	966	7.36	1101

Oxygen (O) + Manganese (Mn)

Foster and Welch, 1956 (fig.)

Atom%	Lattice constant (Å)
45.4	4.355
46.0	.351
46.2	.350
46.5	.349
46.8	.346
47.4	.343

Oxygen (O) + Iron (Fe)

Aubry and Marion, 1956 (fig.)

%	tr.t. 1	2	3
79	900	570	-
76.7	900	"	1350-900
76.4	-	"	570
75.9	-	"	660
75.3	-	"	1350-1000
73.5	-	"	-

1= γ Fe + FeO - α Fe + FeO

2= α Fe + FeO - α Fe + Fe_3O_4 (79-76.4%)

or FeO + Fe_3O_4 - α Fe + Fe_3O_4 (76.4-73.5%)

3= limits of pure FeO

Foster and Welch, 1956 (fig.)

atom%	Lattice constant (Å)
45.4	4.368
46.0	.365
46.5	.360
47.4	.355
47.6	.351
48.5	.366
49.0	.373
49.2	.375
50.5	.375

Chromic trioxide (Cr_2O_3) + Chromium (Cr)

Olchanskaya and Chlepov, 1953

%	f.t.	%	f.t.
100	1900	21.0	1680
99	1810	20.6	1720
29.4	1790	20.0	1660 E
24.7	1680	0	2400

Oxygen (O) + Germanium (Ge)

Trumbore, Thurmond and Kowalchik, 1956 (fig.)

mol%	f.t.		E	
	tetragon.	hexagon.	tetragon.	hexagon.
0	937	-	912	868
16.5	"	-	"	"
28.5	"	-	"	"
37.5	"	-	"	"
44.5	"	-	"	"
46.8	912	-	"	"
47.4	945	868	"	"
48.7	1010	975	"	"
49.3	1035 tr.t.	1035	"	"
50	1085	1115	"	"

Oxygen (O) + Uranium (U)

Hoekstra, Siegel and al., 1955

Phase relationship between uranium oxides (by X-ray study)

Oxygen (O) + Titanium (Ti)

Groves, Hoch and Johnston, 1955

Vapor-solid equilibria by high temperature X-ray diffraction technique

Sulfur (S) + Nickel (Ni)

Sokolova, 1956.

%	f.t.	tr.t.1	E
78.7	700	-	635
76.1	710	-	635
75.7	730	-	-
74.5	766	-	-
73.5	786	-	-
72.2	800	-	-
71.8	804	-	-
71.4	806	-	-
71.0	810	805	-
70.5	816	805	-
69.7	855	805	-
68.7	864	810	-
68.5	896	802	-
68.0	900	800	-
67.2	912	810	-
65.1	975	805	-
64.5	1000	810	-

%	tr.t. 2	3	4	5
78.7	-	-	525 (a)	-
76.1	-	-	525	-
75.7	-	-	530	-
74.5	-	540	525	-
73.5	-	550	525	-
72.2	-	547	518 (b)	-
71.8	-	550	520	-
71.4	-	551	520	-
71.0	-	550	520	-
70.5	-	523	516	-
69.7	570	525	515	400
68.7	575	-	513	400
68.5	578	--	-	400
68.0	575	-	-	410
67.2	570	-	-	400
67.0	585	-	-	400
65.1	570	-	-	405
64.5	577	-	-	380

tr.t. 1 = SNiI - SNi II
2 = S_5Ni_6
3 = S_2Ni_3
4a = S_2Ni_3+Ni
4b = S_2Ni_3+S_5Ni_6
5 = S_8Ni_9

Sulfur (S) + Nickel (Ni)

Sokolova, 1956 (fig.)

atom%	f.t.	atom%	f.t.	tr.t.
100	950	58	790	-
67	700	57.5	805	805
66.5	635 (E)	55.7	860	"
63	730	52.5	900	"
61.5	770	51.0	990	"
60	780	50	995	"

%	d	%	d
	25°		
79.06	6.20	70.50	5.56
78.20	6.20	69.30	5.55
76.33	6.14	68.70	5.49
75.81	6.07	68.40	5.47
74.75	5.94	68.18	5.47
73.50	5.85	68.00	5.48
72.20	5.75	67.20	5.40
71.10	5.61		

%	hardness	%	hardness
	room t.		
79.10	274	71.0	526
76.75	240	70.7	600
75.7	140	70.2	500
73.5	32?	69.3	420
72.6	260	68.28	460
71.9	380	68.0	480
71.7	420	67.8	500

atom%	pressure of flow (kg/mm^2) 550°	620°	(fig.)
69	50	-	
68	52	4	
67.5	53	8	
66	55	12	
63.5	60	22	
62.8	53	20	
61.5	22 (2+3)	16	
57.14	30	13 (3+4)	
56	1100	30	
54.55	23	15 (5+6)	

%	pressure of flow (kg)	%	pressure of flow
	620°		
79.10	6.77	71.10	13.70
78.30	11.00	71.00	14.17
75.70	20.54	70.20	30.40
73.50	19.00	69.30	18.47
72.60	15.13	68.28	20.00
71.80	13.93	68.20	51.75

%	pressure of flow	%	pressure of flow
	550°		
79.10	53.76	70.20	108.00
78.30	54.90	69.30	54.50
73.50	22.50	69.00	41.50
72.60	30.50	68.27	28.00
71.80	26.60	68	55.60
71.00	30.20		

atom%	κ room temp.	107° (fig.)
67.5	5	3.7
64	4.8	3.4
62	4.1	3.0
60	2.8 (2+3)	1.3
58.2	3.3	1.7
56	1.1	0.7
54	0.3	0.2
53	0.4	0.25

%	at room temp.	107°
79.06	5.03	3.65
78.30	4.96	3.44
76.80	4.80	-
76.70	4.72	3.34
76.10	4.63	3.33
75.81	4.54	-
74.75	4.07	2.96
73.50	2.70	-
71.80	3.27	1.77
71.50	2.60	-
71.40	2.27	-
71.20	1.87	-
71.00	1.66	1.27
70.40	1.12	0.94
70.30	1.14	-
70.00	0.88	0.71
69.50	0.65	-
69.20	0.49	-
68.70	0.305	-
68.50	0.407	0.41
68.18	0.36	-
68.00	0.37	0.25
67.20	0.44	0.31

%	κ	%	κ
	380°		
76.14	4.76	71.00	1.75
75.82	4.71	70.30	1.49
74.80	4.08	79.00	0.88
73.50	2.70	78.70	0.61
72.60	2.59	78.50	0.60
71.80	2.95	78.18	0.37
71.50	2.51	76.98	1.01

Sulfur (S) + Bismuth (Bi)

Aten, 1905

atom %	f.t.	m.t.
52.4	277	277
46.9	318	271
40.7	452	272
28.8	535	263
17.2	602	267
4.2	656	270
0.9	702	-
0.0	760	-

Kordes, 1926

mol%	f.t.
100	613
-	267 E

Sulfur (S) + Mercuric chloride ($HgCl_2$)

Olivari, 1909

%	f.t.	tr.t.	
		1	2
100	287	-	-
97.656	293.9	277.5	-
95.057	280.2	-	-
93.546	278	-	-
88.5	-	277.6	117
80	-	277.8	118
70	-	277.2	116.5
50	-	277.5	117
30	-	277.5	117.5
11	-	277	117.5
0	-	-	117

Indium sulfide (In_2S_3) + Indium (In)

Stubbs and al., 1952 (fig.)

%	f.t.	tr.t.				
		1	2	3	4	5
70.5	1090-1100	850	-	-	-	-
72.5	1020	840	750	-	-	-
75	840	840	770	-	640	-
77	770	-	770	680	"	-
78.5	750	-	-	"	"	-
80	680	-	-	"	"	-
85	640	-	-	"	"	155
90	"	-	-	-	"	"
97	"	-	-	-	"	"
99	590	-	-	-	-	"

Selenium (Se) + Mercuric chloride ($HgCl_2$)

Olivari, 1909

%	f.t.	tr.t.	
		1	2
100	287	-	-
97.523	285.3	-	-
95.148	284	-	-
93.283	283.1	-	-
91.075	281.8	272	218.6
86.73	278.9	272.5	-
81.23	275	-	-
69	-	272.5	218
61	-	272.3	218.5
51.8	-	272.2	217.8
42	-	272.5	217.8
30	-	273	217
20	-	272	217
8	-	271.9	217.6
0	-	-	218

Selenium (Se) + Mercuric bromide ($HgBr_2$)

Olivari, 1912

%	f.t.	E	min	tr.t.	min
100	236	-	-	-	-
97.1	233	-	-	-	-
92.7	230.5	227.4	14	210-211	5
90.0	229.8	"	22	"	8
85.6	227.9	"	38	"	10
78.0	-	"	30	"	14
71.8	-	"	27	"	18
65.0	-	"	20	"	21
55.5	-	"	8	"	30
45.0	227.2	-	-	"	48
39.0	226.5	-	-	"	52
32.0	224.4	-	-	"	63
26.0	222.2	-	-	"	80
21.0	?	-	-	"	90
14.0	?	-	-	"	89
8.0	?	-	-	"	90
4.0	?	-	-	"	54
0.0	217	-	-	-	-

Bismuth (Bi) + Tellurium (Te)

Pélabon, 1908

%	f.t.	%	f.t.
0	452	15	410 E
5.92	445	-	583 (3+2)
9.68	437	99	263 E
14.23	417	100	270

Nitrogen (N_2) + Barium (Ba)

Ariya and Prokofieva, 1955

t	P	t	P
	(2+1)		
395.5	35.6	466	140.2
407	46.4	476	169.2
418.5	59.9	486.5	209.4
429.5	75.7	496.5	254.2
442	93.7	507	295.3
454	115.4		

Arsenic (As) + Tin (Sn)

Parravano and De Cesares, 1912

%	f.t.	E	min.
98.07	322	227	300
96.17	395	227	270
94.27	432	229	255
92.37	458	229	240
90.47	478	229	225
88.56	490	231	215
86.63	508	229	210
84.72	518	231	195
82.83	530	231	180
80.93	538	229	165
79.03	554	231	150
77.10	558	229	135
74.61	565	231	105
73.49	571	231	75
71.40	578	231	60
71.97	576	227	-
68.93	583	578	300
66.86	585	578	225
63.20	586	578	80
63.09	588	578	70
62.76	587	-	-
59.87	587	568	45
59.50	585	568	60
56.02	581	568	225

Silicon (Si) + Tungsten (W)

Blanchard and Cueilleron, 1957 (fig.)

atom %	f.t.	atom %	f.t.
0	1420	50	2200
1	1390 E	63	2340 (3+5)
10	1750	70	2290
20	1980	78	2210 E
33	2160 (2+1)	90	2875
41	2010 E	100	3380

Hydrofluoric acid (HF) + Sodium fluoride (NaF)

Jache and Cady, 1952

%	f.t.
23.1	+11.0
20.6	- 9.8
18.0	-24.3

Hydrofluoric acid (HF) + Potassium fluoride (KF)

Cady, 1934

mol %	f.t.	mol %	f.t.
53.5	346	28.6	64.3 (5+2)
51.4	229.5 E	27.3	62.4 E
50.2	236	25.1	65.8 (3+1)
50.1	238 (1+1)	22.4	63.6 E
49.3	234	22.2	67.8 (4+1)
39.9	148	20	72
35.7	84	11.8	8
35.1	68.3 E	8.6	-45
33.3	71.7 (2+1)	6.9	-97 E
30.7	62.4	5.4	-92.8
30.4	61.8 E	0	-83.7
29.6	63.4		

Hydrofluoric acid (HF) + Cesium fluoride (CsF)

Winsor and Cady, 1948

mol %	f.t.	mol %	f.t.
54.7	151.5 E	29.9	16.9 E
50.0	176.0 (1+1)	25.0	32.6 (3+1)
36.1	38.3 E	17.2	49.5 E
33.3	50.2 (2+1)	14.3	42.3 (6+1)

Hydrofluoric acid (HF) + Silver fluoride (AgF)

Jache and Cady, 1952

%	f.t.
45.4	+11.9
30.5	- 9.8
21.3	-25.0

Hydrofluoric acid (HF) + Uranium hexafluoride (UF_6)

Rutledge, Jarry and Davis, 1953

%	f.t.	sat.t.	E
0.0	-83.6	-	-
.16	-5.0	-	-
.24	-5.1	-	-
.27	-5.0	-	-
.32	-5.2	-	-
.41	-5.0	-	-
.45	0.0	-	-
.48	-0.1	-	-
.781	5.0	-	-
.98	25.0	-	-
1.593	26.5	-	-
3.04	44.3	-	-
3.93	50	-	-
4.20	52	-	-
5.39	55	-	-
6.24	55	-	-
6.67	59.12	-	-
6.90	-	-	-84.18
7.95	58	-	-
8.32	60.25	-	-
8.51	59	-	-
9.45	60.55	-	-
10.28	-	69	-
10.38	-	69	-
11.92	60.96	-	-
12.20	-	78	-
16.29	-	83	-
19.85	-	87.9	-
23.39	-	90.3	-
24.52	61.16	-	-
28.58	-	97.2	-
29.13	-	93.2	-
39.38	-	98.8	-
49.97	-	100.5	-
53.14	-	99.9	-
53.53	-	-	-85.13
55.02	61.14	-	-
55.40	-	95.7	-
56.77	61.25	-	-
57.09	-	95	-
58.45	-	98	-
61.35	-	90.5	-
62.03	-	93.5	-
64.49	-	87.5	-
65.21	-	-	85.06
66.70	-	83	-
67.13	-	84	-
69.35	-	79.5	-
70.44	61.18	-	-
72.54	61.25	-	-
72.96	-	76	-
81.56	61.32	-	-
82.03	-	-	85.10
83.06	-	-	84.99
87.73	61.56	-	-
90.46	61.95	-	-
91.09	62.3	-	-
94.15	62.45	-	-
95.06	-	-	85.10
95.77	62.51	-	-
96.47	62.39	-	-
96.47	62.53	-	-
98.45	63.09	-	-
99.63	63.76	-	-
100.00	64.02	-	-

Jarry, Rosen and Dawis, 1953

40.89°

% L	% V	P
0.00	0.00	156.0
2.74	6.85	179.6
100.00	100.00	51.3

50.90°

% L	% V	P
0.00	0.00	214.0
2.47	7.45	244.7
3.81	9.09	258.3
100.00	100.00	73.6

59.66°

% L	% V	P
0.00	0.00	275.0
1.64	7.54	314.4
5.78	11.59	343.8
100.00	100.00	93.9

72.48°

% L	% V	P
0.00	0.00	396.0
1.83	7.38	442.5
2.55	8.00	442.8
4.26	13.13	482.6
7.44	15.04	498.4
10.84	16.15	503.6
73.48	16.40	499.8
77.62	16.93	494.7
93.57	28.40	409.1
98.43	51.15	286.7
100.00	100.00	147.7

66.87°

% L	% V	P
0.00	0.00	340.0
1.89	7.17	382.7
4.07	12.48	412.7
7.40	14.01	429.6
-	13.99	431.1
7.42	13.96	424.5
7.67	15.14	430.5
8.27	14.40	428.3
77.72	15.33	423.1
78.86	14.96	422.6
91.45	24.70	360.8
92.53	25.10	360.1
98.50	46.73	262.0
100.00	100.00	124.5

84.46°

% L	% V	P
0.00	0.00	538.0
1.55	6.07	597.3
3.60	11.60	652.1
7.36	16.00	681.9
11.05	17.75	688.2
11.32	17.81	688.6
65.78	17.80	686.6
73.12	17.56	678.4
93.74	33.33	543.5
98.96	58.75	364.7
100.00	100.00	208.5

92.30°

% L	% V	P
0.00	0.00	656.0
1.50	6.15	718.8
3.50	11.22	789.2
6.91	16.95	830.5
11.24	18.51	840.6
61.67	18.70	837.9
71.17	19.19	821.8
93.94	35.45	637.6
99.53	62.43	421.1
100.00	100.00	257.5

104.74°

% L	% V	P
0.00	0.00	886.0
1.18	5.24	962.4
3.22	10.73	1051.3

104.74°

% L	% V	P
94.92	40.85	804.6
99.56	66.44	524.8
100.00	100.00	352.5

Hydrofluoric acid (HF) + Thallium fluoride (TlF)

Jache and Cady, 1952

%	f.t.
85.2	+11.9
81.8	-7.8
75.3	-25.2

Hydrogen peroxide(H_2O_2) + Sodium fluoride (NaF)

Matheson and Maass, 1929

mol %	f.t.	mol %	f.t.
0	- 1.77	9.8	-12.1
3.8	- 5.2	11.1	-13.9
7.4	- 9.1	12.3	-15.2 E
8.6	-10.6		

Hydrogen peroxide (H_2O_2) + Sodium chloride (NaCl)

Maass and Hatcher, 1922

%	f.t.	%	f.t.
1.08	-2.37	11.81	-10.57
3.14	-3.62	15.63	-10.37
5.42	-5.07	17.00	0
8.61	-7.67		

Hydrogen peroxide (H_2O_2) + Sodium nitrate ($NaNO_3$)

Maass and Hatcher, 1922

%	f.t.	%	f.t.
0	-1.77	16.66	-8.12
1.62	-2.47	20.19	-9.52
3.23	-2.82	22.25	-10.82
5.11	-3.72	25.22	+11.8
7.58	-4.72	28.25	32.3
10.11	-5.62	31.33	49.3
13.19	-6.97		

Hydrogen peroxide (H_2O_2) + Sodium sulfate (Na_2SO_4)

Maass and Hatcher, 1922

%	f.t.	%	f.t.
1.19	-2.17	16.77	-9.02
2.56	-2.87	19.47	-10.57
4.47	-3.12	22.64	+19.6
6.12	-4.27	23.59	+29.1
10.23	-5.52	25.62	+39.1
13.51	-6.77		

Hydrogen peroxide (H_2O_2) + Potassium chloride (KCl)

Matheson and Maass, 1929

mol%	f.t.	mol%	f.t.
2.6	-3.4	15.9	-31.4
4.8	-6.5	17.0	-24.2
7.1	-10.4	18.7	-16.0
9.4	-14.6	20.5	-7.0
11.9	-20.2	22.1	-1.0
13.9	-25.0		

Hydrogen peroxide (H_2O_2) + Potassium sulfate (K_2SO_4)

Matheson and Maass, 1929

mol%	f.t.	mol%	f.t.
0.87	-2.0	7.6	-17.5
2.27	-4.8	8.4	-23.5
3.68	-8.2	14.1	-11.0
5.3	-12.3	15.9	+0.5

Heavy water (D_2O) + Sodium chloride (NaCl)

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
9.06	-3.8	9.82	61.5
9.14	1.7	10.5	110.8
9.29	11.7	11.0	144.1
9.36	19.7	11.7	180.3
9.47	28.2		

Heavy water (D_2O) + Sodium bromide (NaBr)

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
11.9	-2.3(2+1)	16.4	35.0
12.5	10.1	16.4	50.2
13.1	23.4	16.5	57.9
15.0	35.4	16.6	75.7
15.7	39.95	17.1	99.7
16.0	44.0	17.3	122.9
16.2	46.6	17.9	132.9
16.4	47.5	18.7	164.9

Heavy water (D_2O) + Sodium iodide (NaI)

Eddy and Menzies,1940

mol%	f.t.	mol%	f.t.
16.6	12.4(2+1)	24.2	46.45(x+1)me-
18.2	27.78	24.6	51.7 tast.
20.4	42.50	24.9	56.9
22.1	54.04		
23.7	60.45	25.7	62.1NaI metast.
24.6	64.02	25.9	77.7
24.6	64.07	26.5	112.2
24.9	65.44		
25.5	65.55		

Heavy water (D_2O) + Sodium sulfate (Na_2SO_4)

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
2.28	8.7(10+1)	3.51	7.7(7+1)
2.41	11.5	3.74	10.3
2.76	16.2	4.00	14.3
3.22	22.4	4.17	17.9
3.74	26.2	5.55	24.7
4.00	29.0	5.74	27.0
4.33	30.55		
		5.74	24.8 anh.
		5.55	42.8
		4.29	124.2
		4.30	152.0
		4.33	192.7

Heavy water (D_2O) + Potassium bromide (KBr)

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
6.9	8.0	12.0	82.7
8.0	12.2	13.4	96.2
8.6	20.1	14.4	126.9
8.8	31.10	16.2	167.2
9.4	48.98	19.1	203.8
11.7	67.7		

Heavy water (D_2O) + Potassium iodide (KI)

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
10.75	2.6	15.25	62.7
11.50	10.7	16.85	88.6
12.11	17.8	18.63	118.1
13.34	34.0	20.19	146.1
15.08	60.3		

Heavy water (D_2O) + Magnesium chloride ($MgCl_2$)

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
1.03	8.1	1.91	57.4
1.22	18.7	2.33	67.3
1.51	27.2	2.71	91.2
1.72	48.9	3.51	105.1

Heavy water (D_2O) + Barium chloride ($BaCl_2$)

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
2.71	4.1(2+1)	3.64	95.5(1+1)
3.47	21.2	3.98	104.8
3.48	21.6	4.12	119.8
3.53	37.0	4.26	135.0
3.76	83.1	4.40	159.3
4.37	77.1	5.36	179.7
3.54	80.7		

Heavy water (D_2O) + Strontium chloride ($SrCl_2$)

Miles and Menzies, 1937

mol%	f.t.	mol%	f.t.
4.33	5.83(6+1)	9.32	93.5
5.51	20.62	9.43	96.45
5.74	31.32	11.68	116.34
8.61	50.75	11.85	125.2
8.73	55.83	11.93	127.8
8.82	57.37	12.21	133.8
8.82	60.76(2+1)	12.26	137.3
8.91	70.6	13.0	141.0
8.96	79.2		

Heavy water (D_2O) + Cadmium iodide (CdI_2)

Eddy and Menzies, 1940

mol%	f.t.	mol%	f.t.
3.47	1.7	4.38	96.9
3.50	11.1	4.97	115.0
3.63	24.4	5.24	125.9
3.73	44.4	5.50	137.4
4.00	57.7	6.60	149.4
4.11	78.9		

Heavy water (D_2O) + Cupric sulfate ($CuSO_4$)

Miles and Menzies, 1937

mol%	f.t.	mol%	f.t.
2.9	2.4	4.1	63.2
3.2	1.5	5.2	79.2
		6.9	90.3
2.9	11.6(5+1)	8.2	99.55
3.2	28.4		
3.7	47.75	8.5	104.7(3+1)

Ammonia (NH_3) + Lithium (Li)

Kraus and Johnson, 1925

mol%	p	mol%	p
	-39.4°		
36.8	3.5	7.077	532.2
22.7	3.7	6.82	541.3
21.0	45.2	6.220	534.1
20.9	47.5	6.05	541.4
20.8	51.8	5.516	536.8
19.7	134.8	4.990	536.2
18.6	174.8	4.87	544.5
17.1	290.6	4.478	538.7
17.0	281.7	4.085	545.0
14.7	395.5	4.06	542.5
13.1	443.2	3.732	538.8
11.7	474.8	3.143	543.6
10.9	492.8	2.712	540.3
9.67	506.7	2.143	539.7
9.14	518.4	1.778	546.8
8.354	520.3	1.636	549.6
7.84	533.4		

Marshall and Hunt, 1956

mol %	p	mol %	p
	-63.8°		
1.72	126.35	7.04	123.85
1.85	125.90	10.68	117.10
1.95	125.45	15.80	83.90
2.07	124.95	20.24	15.65
2.24	124.50	20.88	11.65
2.43	124.70	22.08	0.05
2.66	124.35	27.25	0.20
3.04	124.95	32.57	0.05
3.50	124.80	37.31	0.10
4.21	124.40	40.00	0.05
5.21	124.50	41.49	0.10
	-45.3°		
0.146	399.80	2.21	392.80
0.180	399.50	3.09	391.45
0.232	399.35	3.70	390.90
0.332	399.00	6.10	385.85
0.422	398.70	7.63	381.55
0.599	398.35	8.55	372.75
0.649	398.50	11.49	338.55
0.685	399.60	14.7	268.55
0.709	398.45	18.1	148.40
0.714	398.50	19.0	113.60
0.752	397.45	18.4	84.80
0.800	396.90	20.4	15.85
1.094	396.40	20.9	0.55
1.34	395.70	25.6	0.50
2.08	393.30	39.5	0.55

t	p	t	p
V+sat.sol. + C_2		V + L_1 + L_2	
- 63.8	0.05	- 63.8	124.50
- 45.4	0.55	- 45.4	390.90
- 22.9	7.60		
0.0	32.85		
21.3	103.9		

Johnson, Meyer and Martens, 1950

mol%	d	mol%	d
	-33.2°		
3.31	0.639	15.57	0.523
5.71	.611	16.16	.518
7.13	.597	19.58	.498
11.80	.554	21.05	.490

Birch and Mc Donald, 1948 (fig.)

mol%	f.t.	mol%	f.t.
5	-80	19	-91
11	-82	20	-100
15	-85	21	-110
18	-89	22	-186

Marshall and Hunt, 1955

mol %	d	Dv (cc/mole)
3.31	0.639	46.7
5.71	0.611	46.1
7.13	0.597	44.8
11.80	0.554	42.5
15.57	0.523	41.4
16.16	0.518	41.4
19.58	0.498	38.8
21.05	0.490	37.9

Ammonia (NH_3) + Lithium chloride (LiCl)

Bonnefoi, 1898

t	p dissoc.	
119	975	(1+1)-LiCl
109.2	646	
96	367	
88	256	
89.2	980	(2+1)-(1+1)
83	739	
77	558	
68.8	373	
65	1011	(3+1)-(2+1)
62.2	882	
60	790	
50	473	
43	320	
32.8	177	
14.5	800	
13.74	754	
9	640	(4+1)-(2+1)
0	384	

Ammonia (NH_3) + Lithium bromide (LiBr)

Bonnefoi, 1901

(1+1)		(2+1)	
t	p dissoc.	t	p dissoc.
368.5	730	358	655
370.5	806	363.2	847

(3+1)		(4+1)	
t	p dissoc.	t	p dissoc.
340	636	329	681
344.8	807	331.25	760

Ammonia (NH_3) + Lithium nitrate ($LiNO_3$)

Kamejáma, 1931

t	p	t	p
	sat.sol.		
10	102	30	230
20	155	50	455

Davis, Olmstead and Lundstrum, 1921

t	p	t	p
	63.66%		
-13.0	239.32	15.5	487.45
-6.02	259.94	22.4	627.39
+1.53	321.90	30.05	818.13
8.48	382.90	35.0	998.16

Portnov and Dwilewitch, 1937

%	f.t.	%	f.t.	%	f.t.
0	-77.0	36.70	-31.0	66.38	+14.5
3.03	-77.8	40.51	-28.0	69.14	12.5
10.73	-79.0	42.97	-12.0	70.00	25.0
21.74	-80.5	49.28	+ 5.5	71.59	42.0
24.26	-78.5	52.82	+ 3.0	78.35	119.0
30.62	-62.0	56.96	- 3.0	82.60	160.0
33.51	-53.0	60.86	+ 3.0	89.61	204.0
35.66	-54.0	62.70	+ 7.0	100.00	264.0
(8+1)	(4+1)	(2+1)			

%	d	%	d
		20°	
0	0.610	35.82	0.890
16.41	0.742	45.44	0.994
21.72	0.772	62.12	1.173
31.80	0.863		

Ammonia (NH_3) + Sodium (Na)

Birch and Mc Donald, 1947

C.S.T. = -74° E = -110°

Joannis, 1890

t	p dissoc.	
0	169.7	
-10	117.0-117.3	(1+1)
+22.4	371.2-371.8	

Kraus, Carney and Johnson, 1927

mol %	p	mol %	p	mol %	p
		-33°			
17.17	405.0	12.99	577.5	16.40	403.0
15.57	398.0?	12.13	623.6	15.48	403.0
15.23	421.0	11.35	656.6	15.15	431.5
14.90	445.0	10.07	686.5	14.27	486.0
14.76	461.0	9.505	707.5	13.39	544.5
14.56	469.0	8.5143	745.5	12.64	592.5
14.43	483.5	6.8677	769.0	12.03	625.0
14.18	493.0	26.05	397.0	11.21	661.5
13.99	516.0	20.02	397.0?	10.68	681.5
13.71	533.5	18.08	401.0		

Marshall and Hunt, 1956

mol%	p	mol%	p	mol%	p
-63.8°		-45.3°		-45.3°	
1.33	126.85	0.144	399.85	2.40	395.10
.35	126.45	.162	399.70	2.74	394.55
.40	126.25	.153	399.55	2.75	394.50
.48	126.30	.214	399.45	3.00	395.25
.59	126.25	.253	399.30	3.37	394.45
.87	126.35	.306	399.10	3.95	394.50
2.30	126.40	.391	398.90	4.46	394.55
3.18	126.45	.526	398.55	5.52	393.60
4.83	126.35	.610	398.45	7.58	388.60
7.46	126.25	.685	398.10	9.91	370.40
10.4	121.55	.735	397.80	11.6	330.70
12.5	108.00	.794	397.70	13.1	295.35
13.8	95.05	.962	397.45	13.7	270.10
14.6	85.40	1.18	396.90	14.6	245.10
15.4	76.60	.25	396.45	15.1	227.25
15.9	66.25	.53	396.10	15.6	206.70
17.5	63.40	.56	395.90	15.7	200.10
26.8	63.35	.64	395.75	15.8	207.55
		.79	395.40	20.6	200.10
		2.03	395.35	35.3	200.15
		.12	395.10		

t	p	t	p
V+sat.sol.+C		L_1+L_2+V	
-63.8	63.35	-63.8	126.30
-45.4	200.10	-45.4	394.50
-22.9	637.50		
0.0	1697.25		

d	N	d	N
	-33.8°		
0.5782	4.96	0.6322	2.23
.5888	4.27	.6376	1.99
.6044	3.47	.6423	.79
.6163	2.91	.6462	.64
.6251	2.54	.6494	.49

N	κ	N	κ
4.96	5047	4.29	3166
.78	4348	.07	2687
.78	4394	3.88	2347
.51	3718	.63	1945

Ruff and Zedner, 1908

%	b.t.	%	sat.t.
1.23	-33.33	0.68	miscible
1.90	-33.27	0.81	"
3.62	-33.20	1.08	"
5.25	-33.15	1.88	-50
6.95	-33.1	2.69	-45
8.43	-33.0	3.75	-47
10.70	-32.6	5.33	-49
14.53	-30.5	7.82	-54
18.03	-27.0	9.89	-60
21.13	-19.4	10.92	-61
23.26	-19.4	13.58	-65
		16.70	-72

%	f.t. L_1	f.t. L_2	min. L_1	min. L_2
0	-76	-	17.2	-
3.09	-77	-115	13.3	4.3
5.33	"	-114	10.6	7.4
9.12	"	-111	8.0	11.3
10.92	-78	-109	6.0	12.7
13.58	-79	"	4.8	15.2
17.07	"	"	3.7	-
19.28	"	-111	2.0	-
20.38	-81	"	1.0	-
22.31	-	"	-	-

Kraus and Lucasse, 1922

mol%	f.t.	mol%	f.t.
7.09	-51.6	2.87	-44.6
4.46	-41.8	2.64	-46.4
4.97	-42.5	4.18	-41.6
5.47	-44.4	2.50	-47.5
6.27	-47.4	2.38	-48.6
8.34	-60.0	2.09	-51.9
3.83	-41.7	1.81	-55.9
3.18	-42.8	1.62	-59.7

Kraus, Carney and Johnson, 1927

atom %	d	atom %	d
	-33.8°		
15.43	0.5782	5.43	0.6376
12.93	.5888	4.85	..6423
10.14	.6044	4.39	.6462
8.31	.6163	3.99	.6494
7.09	.6251	3.54	.6535
6.15	.6322	2.98	.6582

Johnson and Meyer, 1932

mol %	d	mol %	d
	-33.8°		
sat.t.	0.5781	4.4084	0.6681
6.3000	.6603	3.8153	.6704
5.0297	.6653		

Marshall and Hunt, 1955 (fig.)

N	Dv cc/atom-gr.	N	Dv cc/atom-gr.
1.2	41.5	4.0	42.5
2.0	43	5.0	41
3.0	43.5		

Kraus and Lucasse, 1921

mol/l solvent	ϰ (in mhos)	mol/l solvent	ϰ (in mhos)
	-33.5°		
9.201	5047	3.115	726
8.333	5047	3.096	714.4
7.513	4954	2.947	612.8
6.944	4348	2.888	575.3
6.925	4394	2.493	379.9
6.410	3718	2.346	297.6
6.188	3167	2.276	275.8
6.150	3356	1.961	148.3
5.949	3166	1.733	92.19
5.192	2347	1.685	79.56
4.897	1981	1.329	22.95
4.764	1945	1.314	20.21
4.049	1344	1.080	6.493
3.984	1288	1.079	5.988
3.771	1174	0.7704	1.269
3.622	1080	0.1195	0.6456
3.613	1070		

Kraus and Lucasse, 1922

mol/l solvent	τ *	mol/l solvent	τ *
	between -32° and -50°		
7.097	0.0656	0.6636	2.89
5.663	0.158	0.6250	2.57
4.316	0.162	0.6024	2.54
4.055	0.290	0.5522	2.41
3.511	0.369	0.4965	2.19
3.141	0.447	0.4677	2.11
2.559	0.633	0.4560	2.10
2.203	0.807	0.4160	1.77
1.879	1.000	0.3811	1.85
1.524	1.45	0.3794	1.84
1.300	2.06	0.3672	1.17
1.26	3.04	0.3597	1.57
1.076	3.35	0.3319	1.57
1.06	3.38	0.3258	1.57
1.02	3.50	0.2958	1.63
0.9074	3.57	0.2685	1.57
0.8749	3.60	0.2317	1.59
0.8354	3.60	0.2009	0.56
0.7746	3.27	0.1879	0.61
0.7112	3.05	0.1795	0.52
		0.1694	0.52

mol/l solvent	t_1	t_2	t_3	τ * t_1-t_2	τ * t_2-t_3
7.097	-32.2	-50.0	-60.0	0.0656	0.0930
5.662	-32.3	-50.1	-61.0	.158	.183
5.663	61.0	68.0	-	3.44	-
4.316	-32.2	-40.0	-51.0	0.262	0.318
0.5875	-32.5	-40.0	-50.0	2.11	2.24
0.5875	50.0	60.0	-	0.201	-
0.4677	-32.3	-40.0	-50.0	2.11	2.24
0.4677	50.0	59.0	68.0	2.64	3.08

*τ = resistance-temp. coeff. $1/R_{-33.5} \cdot (DR/Dt).100$

Dewald and Lepoutre, 1954

m	th. power*	m	th. power*
	33°		
1.026	78.2	6.40	4.08
1.321	57.0	7.72	2.68
1.702	29.8	7.87	2.50
2.15	16.6	8.33	1.995
2.81	9.95	9.40	1.11
3.335	8.18	9.68	0.65
3.38	6.78	10.47	-0.225
4.58	5.67	10.70	-0.305
4.66	5.67	10.72	-0.43
5.49	4.91		

* = thermoelectric power (mv./°c)

The sign of the Seebeck voltage is taken as positive when electrons tend to flow from hot to cold in the solution .

Ammonia (NH_3) + Sodium chloride (NaCl)

Joannis, 1891

t	p dissoc.	t	p dissoc.
	(5+1)		
-24	777	-15.0	1305
-20.8	892	-10.0	1777
-17.5	1074	-7.0	2130

Abe and Hara, 1933

%	t	P	
3.221	30	11.402	
4.118	25	9.783	
5.121	20	8.337	
6.630	15	7.066	
8.324	10	5.944	
9.500	5	4.962	
11.316	0	4.101	
13.284	- 5	3.364	
14.733	- 8.5	2.912	
14.851	-10	2.737	(5+1)
14.080	-11	2.637	
13.617	-11.5	2.588	
12.316	-13	2.443	
9.613	-18	1.998	
6.628	-23	1.611	

Linhard, 1936

m	p	m	p
	0°		
0	3220.9	1.378	3156
0.178	3210	1.865	3133
1.071	3170	2.200	3118

Abe and Hara, 1933

%	f.t.	%	f.t.
3.221	30	14.113	-7
4.118	25	14.733	-8.5
5.121	20	14.851	-10 (5+1)
6.630	15	14.080	-11
8.324	10	13.617	-11.5
9.500	5	12.316	-13
11.316	0	9.613	-18
13.284	-5	6.628	-23

Patscheke, 1933

%	f.t.	%	f.t.
0.000	-76.3	11.49	-14.2
.074	-76.8	12.20	-13.1
.132	-76.5	12.65	-12.5
.232	-76.9	13.83	-11.2
.28	-76.6	14.65	-9.9
.306(5+1)	-75.1	15.37 E	-9.5
.428	-68.1	16.27	-8.5
.465	-66.9	16.26(NaCl)	-11.5
.521	-64.5	13.84	-6.0
.629	-59.5	12.65	-3.4
.945	-52.5	12.20	-2.6
1.763	-43.0	11.52	0
4.224	-28.8	15.55	+2.1
6.206	-23.2	9.609	+4.5
6.923	-22.0	8.600	+7.2
7.208	-20.0	8.506	+8.0
7.54	-20.7	8.44	+7.2
8.44	-18.6	7.54	+11.0
8.506	-18.6	7.208	+12.3
8.600	-18.6	6.923	+13.0
9.609	-16.7	6.206	+15.9
9.871	-16.5	4.224	+24.9
10.28	-15.8	2.970	+32.7
10.55	-15.2	1.763	+43.0

Abe, Sigetom and Hara, 1935

%	f.t.	%	f.t.
3.30	30	16.73	-10
4.11	25	15.81	-11
4.38	20	14.83	-11.5
6.21	15	13.22	-13
8.48	10	9.01	-18
10.42	5	6.03	-23
12.01	0	3.72	-30
14.38	-5	1.62	-40
16.00	-7	0.61	-50
16.54	-8.5	0.22	-60

Kikuti, 1939

%	d					
	-30°	-20°	-10°	0°	10°	20°
0.0	0.6776	0.6650	0.6520	0.6386	0.6247	0.6103
2.5	.6939	.6815	.6687	.6555	.6417	.6278
4.22	.7049	-	-	-	-	-
5.0	-	0.6976	0.6849	0.6719	0.6584	0.6445
5.45	-	-	-	-	-	.6476
7.5	-	0.7132	0.7007	0.6878	0.6743	-
8.55	-	.7197	-	-	-	-
8.59	-	-	-	-	0.6817	-
10.0	-	-	0.7160	0.7033	-	-
12.5	-	-	.7311	.7185	-	-
12.83	-	-	-	.7205	-	-
15.0	-	-	0.7459	-	-	-
17.47	-	-	.7604	-	-	-

Ammonia (NH_3) + Sodium bromide (NaBr)

Linhard, 1936

molality	p	molality	p
	0°		
0	3220.9	1.920	3114
0.311	3203	3.602	2987
0.892	3173	5.013	2835
1.642	3131	6.21	2661

Scherer, Jr., 1931

c	f.t.	c	f.t.
7.37	-44.6	10.32	-37.2
8.03	-42.6	10.56	-36.6
8.34	-42.3	10.64	-36.8
9.09	-40.5	10.88	-36.3
9.64	-38.7	11.24	-35.2
9.62	-39.3	11.53	-35.3
10.23	-37.5		

Distanov, 1937

%	f.t.	%	f.t.	
25	-22	53.76	12.8	
31.75	-14	54	13	
36	-8	54	75	
41.06	-2	53.76	88	(5+1)
44.5	2	52.8	107	
46.1	2.5	51.8	120	
48.96	10	51.07	129	
50.1	10.5	50.1	134	
50.35	10.9	48.8	147	
51.8	11.9	47.2	156	
52.8	11.5	41.00	160	

Johnson and Meyer, 1932

mol%	d	mol%	d
	-33.2°		
2.174	0.7417	0.6994	0.7073
1.493	.7273	.4313	.6983
1.145	.7204	.3152	.6941
0.9689	.7153	.2848	.6933

Ammonia (NH_3) + Sodium iodide (NaI)

Linhard, 1936

molality	p	molality	p
	0°		
0	3220.9	3.426	2938
0.174	3209	5.657	2574
0.655	3181	7.030	2243
1.038	3161	8.324	1869
1.833	3098	8.800	1862-3

Scherer, Jr., 1931

%	f.t.	%	f.t.
44.8	-42.2	49.0	-37.4
45.5	-41.9	51.3	-35.2
48.4	-37.5	54.6	-31.5
48.8	-37.8		

Ammonia (NH_3) + Sodium nitrate ($NaNO_3$)

Kameyama and Yagi, 1928

t	p	t	p
	sat.sol.		
-35.0	325	-16.7	760
-30.7	400	-15.6	800
-26.0	500	-10.5	1000
-22.1	600	-6.3	1200

Schattenstein and Uskowa, 1935

molality	p	molality	p
	15°		
0.222	5443.5	1.160	5362.5
.229	5440.5	.744	5309.5
.454	5421.5	2.267	5262.5
.646	5405.5	.493	5233.5
.967	5378.5	.525	5229.5
.981	5378.5		

Linhard, 1936

m	p	m	p
	0°		
0	3220.9	7.156	2594
0.188	3210	9.599	2262
0.400	3198	11.96	1949
0.778	3179	13.34	1782
1.599	3134	14.35	1675
3.067	3036	14.43	1666
5.064	2853	15.00	1605

Franklin and Kraus, 1898

%	D.b.t.	%	D.b.t.
1.51	+0.043	14.30	+0.731
1.62	.054	20.01	1.190
4.09	.150	24.87	1.782
8.84	.384	26.06	1.997
10.01	.457	29.25	2.547

Scherer, Jr., 1931

%	f.t.	%	f.t.
47.3	-50.5	58.5	-48.0
50.3	-49.0	59.0	-48.2
52.3	-48.0	59.9	-50.2
53.3	-47.5	61.2	-48.6
55.8	-47.6	63.1	-49.8

Fitzgerald, 1912

N	d	N	d
	-33.5°		
3.295	0.896	0.1816	0.6935
1.534	.788	.1121	.687
0.9017	.746	.1065	.6895
.9000	.746	.08475	.687
.5280	.718	.06246	.685
.4202	.710	.04973	.684
.3097	.702	.03664	.683
.2464	.697	0	.6823

Schattenstein and Uskowa, 1935

%	d
15°	
1.62	0.628
4.03	.643
10.17	.683
17.18	.733
18.04	.739
21.40	.757
22.60	.781

Fitzgerald, 1912

N	η	N	η
	-33.5°		
3.295	878.7	0.1816	280.8
1.534	446.4	0.1121	270.5
0.9017	364.4	0.1065	275.4
0.9000	360.9	0.08475	272.4
0.3097	293.0	0.06246	271.6
0.4202	304.4	0.04973	269.0
0.3097	293.0	0.03664	269.0
0.2464	287.9	0	265.6

Akhumov, Goncharov and Ezerova, 1935

t	κ
sat.sol.	
-16	768
-15	768
- 7	1107
0	1269
+ 3	1400
+10	1565
+15	1645-53
+19	1716-06
+30	1776

Ammonia (NH_3) + Potassium (K)

Johnson and Meyer, 1929

mol%	p	mol%	p
	0°		
46.62	748.3	15.60	1140.2
28.23	755.9	15.34	1161.5
24.25	747.1	15.12	1229.9
23.47	760.8	14.94	1226.9
21.95	764.4	14.63	1356.1
20.20	753.3	14.56	1398.3
19.88	748.9	14.54	1393.6
19.78	799.5	14.53	1354.9
19.27	750.4	14.29	1405.7
18.02	748.9	14.10	1501.5
17.98	760.1	14.06	1491.7
17.81	741.9	13.84	1651.6
17.51	792.6	13.48	1632.0
17.36	769.9	13.08	1740.9
17.24	832.8	12.69	1854.0
17.22	835.5	12.44	1901.9
16.79	884.8	12.29	1927.6
16.53	937.4	12.28	1975.1
16.44	965.2	12.19	2001.4
16.27	992.9	11.97	2062.2
16.39	918.3	11.81	2018.8
16.02	1005.4	11.89	2018.2
15.79	1082.5	11.04	2161.2
15.76	1091.2	10.42	2360.5
15.64	1070.7		

mol%	p	mol%	p
	-33.5°		
37.64	165.0	13.82	396.2
34.64	159.7	13.64	351.7
31.91	162.1	13.20	675.2
29.53	162.0	13.19	393.6
26.85	159.5	12.85	405.0
26.70	163.5	12.22	500.3
21.57	161.6	11.07	566.9
21.30	161.0	10.78	564.9
20.95	161.0	10.69	561.3
20.41	167.0	10.60	571.4
18.42	162.8	10.60	583.5
17.94	166.2	10.37	599.5
17.01	162.0	10.31	601.2
16.95	161.0	10.18	640.5
16.82	162.2	9.254	640.5
16.71	172.8	8.906	657.4
16.64	166.0	8.744	669.8
16.48	168.7	8.728	666.3
16.44	177.2	8.143	686.0
16.19	191.0	7.635	693.6
16.16	202.8	7.564	763.0
16.00	193.7	7.4239	707.2
15.94	208.0	7.250	711.8
15.48	226.4	7.108	714.0
15.38	294.5	7.1037	722.5
14.89	265.1	6.677	722.2
14.27	309.7	6.1610	732.3
13.90	359.6	4.980	745.0

mol%	p	mol%	p
	-50.38°		
40.24	62.9	15.95	83.6
37.78	63.4	15.83	81.2
22.78	63.5	14.01	145.2
21.34	61.8	13.79	152.3
16.86	62.6	13.56	159.1
16.71	64.8	13.34	166.4
16.65	62.6	13.11	174.4
16.57	63.6	12.67	183.9
16.50	67.5	12.44	187.9
16.43	64.6	12.30	193.2
16.31	72.0	10.41	236.6
16.29	68.7	10.20	240.9
16.14	72.0	8.878	258.0
16.13	76.3	8.374	276.5
15.98	76.6	7.752	278.3

Ruff and Zedner, 1908

atom %	b.t.	atom %	b.t.
0.76	-33.2	14.33	-18.3
1.49	-33.1	17.32	-2.4
2.36	-33.0	17.66	-1.2
3.52	-32.7	18.28	-0.6
5.54	-32.5	18.87	-0.4
10.50	-29.0	20.02	-0.2 sat.sol.
11.73	-26.6	21.64	-0.2 " "
12.88	-23.9		

atom %	f.t.	min.	atom %	f.t.	min.
0.0	-76	13.2	7.0	-79	7.2
2.0	-77	12.5	9.7	-79	6.0
4.1	-77	10.2	12.0	-83	2.8
5.9	-78	9.8			

Johnson and Meyer, 1929

mol%	f.t.
17.60	0
16.81	-33.5
16.53	-50.0

Birch and Mc Donald, 1948

mol%	f.t.	mol%	f.t.
6	-70	14	-100
11	-75	14.5	-103
11.5	-76	15	-158

Johnson and Meyer, 1932

mol%	d	mol%	d
	-33.2°		
16.8	0.6282	5.333	0.6598
12.45	.6351	4.136	.6659
9.39	.6433	3.602	.6685
8.197	.6483	3.161	.6701
7.042	.6522	2.851	.6721
5.807	.6583		

Marshall and Hunt, 1956

N	d	N	d
	-33.2°		
5.09	0.6282	2.06	0.6583
4.00	.6351	1.93	.6598
3.16	.6433	1.54	.6659
2.82	.6483	1.35	.6685
2.47	.6522		

N	Dv cc/atom-gr.
1.2	25.5
2.0	28.5
3.0	29.5
3.5	29.5
4.0	29.0
5.0	27.5

Kraus and Lucasse, 1921

mol/L solvent	κ (in mhos)	mol/L solvent	κ (in mhos)
	- 33.5°		
11.72	4569	4.136	1288
9.452	4560	4.055	1278
7.722	4190	3.168	701.4
6.583	3233	3.015	640.7
6.361	2940	2.570	389.4
5.780	2571	2.384	318.8
4.990	1858	1.793	91.116
4.755	1772	1.383	21.69

Kraus and Lucasse, 1922

atom/L solvent	τ*	atom/L solvent	τ*
	-33.5°		
8.000	0.0438	1.014	4.41
7.710	0.0880	0.9268	4.16
6.258	0.132	0.8569	3.91
4.888=	0.208	0.7097	3.04
3.532	0.388	0.5230	4.55
2.805	0.613	0.5322	2.22
2.489	0.806	0.4093	1.85
2.281	1.02	0.3184	1.67
1.923	1.57	0.2576	1.56
1.782	1.90	0.2489	1.49
1.538	2.80	0.1946	1.46
1.400	3.43	0.2218	1.47
1.242	4.22	0.1892	1.48
1.123	4.50	1.1644	1.42
1.057	4.50		

*τ = Resistivity temperature coefficient : $1/R_{-33.5} \cdot (DR/Dt) \cdot 1000$

Marshall and Hunt, 1956

N	κ (in mhos)	N	κ (in mhos)
	-33.2°		
5.09	4569	4.05	9571
4.93	4190	3.64	1858
4.93	3233	3.51	1772
4.33	2940		

Dewald and Lepoutre, 1954

m	thermoelectric* power (mv./°c)
-33°	
2.30	20.1
2.85	12.4
4.46	4.86
5.49	3.99
6.46	3.40
7.41	2.86
7.62	2.75
8.23	2.40
8.61	2.12
9.12	1.89
9.29	1.22
9.65	1.60
9.87	1.48
10.41	1.24
12.20	0.91

*-the sign of the Seebeck voltage is taken as positive when electrons tend to flow from hot to cold in the solution.

Ammonia (NH_3) + Potassium bromide (KBr)

Linhard, 1936

m	p	m	p
	0°		
0	3220.9	1.558	3139
0.210	3209	1.962	3117
0.582	3190	2.26	3100
1.097	3163		

Ammonia (NH_3) + Potassium iodide (KI)

Linhard, 1936

m	p	m	p
	0°		
0	3220.9	4.543	2776
0.242	3206	5.680	2585
0.469	3192	6.962	2332
1.014	3158	9.196	1855
2.674	3015	10.24	1642
2.992	2981	11.09	1474

Franklin and Kraus, 1898

%	D.b.t.	%	D.b.t.
36.06	+1.994	14.48	+0.408
30.30	.351	14.43	.412
29.64	.678	12.92	.346
27.42	.129	11.97	.315
21.23	0.720	11.76	.306
20.14	.650	11.75	.306
17.19	.526	10.18	.255
17.02	.507	9.92	.248
16.49	.480	7.39	.178
16.21	.471	6.23	.155
15.54	.491	5.40	.109
15.15	.433	2.12	.038
15.11	.441	1.57	.016

Fitzgerald, 1912

N	d	N	d
	-33.5°		
4.434	1.2110	0.8230	0.8640
2.124	0.9790	.6006	.7671
1.718	1.054	.4474	.7814
1.155	0.8441	.2876	.7243
1.097	0.8385	.1564	.7053

N	d	η
4.312	1.275	1446
2.008	0.963	579.9
1.170	.850	432.6
0.5450	.755	313.2
.3195	.7275	287.1
.1874	.709	281.1
.1099	.6965	273.0
.06447	.689	268.2

Ammonia (NH_3) + Rubidium bromide (RbBr)

Linhard, 1936

m	p
0°	
0.237	3207
0.530	3103
0.979	3173
1.35	3159
0	3220.9

Ammonia (NH_3) + Rubidium iodide (RbI)

Linhard, 1936

m	p	m	p
	0°		
0	3220.9	4.404	2840
0.174	3210	6.649	2513
.407	3196	7.759	2326
.901	3166	9.139	2097
1.801	3104	9.589	2025
2.792	3019	10.08	1952

Ammonia (NH_3) + Cesium (Cs)

Hodgins, 1949

mol%	p	mol%	p
-50°			
36.2	43.5	12.0	203.4
31.1	43.8	11.0	224.6
28.0	44.9	9.77	245.7
25.6	46.3	9.27	252.6
24.3	47.6	8.85	257.2
23.2	51.0	8.92	258.0
22.1	56.2	8.65	260.0
21.1	59.0	8.45	264.7
20.2	63.8	8.10	268.3
19.4	72.7	7.75	273.5
18.7	80.0	7.46	277.0
18.1	83.3	7.17	284.6
18.0	84.9	6.66	282.8
17.0	97.4	6.66	284.0
16.1	113.1	6.43	285.8
15.3	128.6	6.22	286.8
14.5	146.9	5.40	291.8
13.2	177.0		

mol%	d	mol%	d
-50°			
14.0	1.0140	5.211	0.8570
9.80	0.9545	4.435	.8392
7.576	.9095	3.165	.8041
6.154	.8788		

mol%	κ (in mhos)	mol%	κ (in mhos)
-70°			
11.3	650.0	2.99	0.81
8.06	400.0	2.70	.63
6.28	222.0	1.78	.332
5.13	89.6	0.892	.133
4.35	23.8	.623	.084
3.77	3.12	.102	.016
3.34	1.36		

Ammonia (NH_3) + Cesium iodide (CsI)

Linhard, 1936

m	p	m	p
0°			
0	3220.9	2.349	3085
0.185	3210	4.382	2922
.616	3186	5.342	2828
.941	3169	5.84	2773
1.478	3139		

Ammonia (NH_3) + Cuprous chloride (CuCl)

Lloyd, 1908

mol%	p dissoc. 90°	t	p dissoc. (3+1)
14.08	783(3+1)	-	-
15.12	780	88	736
16.39	779	93	915
18.51	781	98	1132
22.73	778	110	1650
25.64	720	-	-
27.03	660	-	-
28.57	650(3+2)	67.5	324
33.33	651	85.0	567
40.00	648	92.0	700
47.62	649	100.0	845
52.63	620	87.0 (1+2)	518
55.56	530	90.0	523
60.68	522	120.0	680
68.97	523	135.5	819
76.92	522	-	-
83.33	312	-	-
90.91	40	-	-

Peters, 1912

t	p dissoc. (3+1)
30	334-340
25	263-264
20	196-199
17.1	167- -
16.7	- -170
15	149-153
10	112-114
5	85- 88

Ammonia (NH_3) + Cuprous bromide (CuBr)

Lloyd, 1908

mol%	p dissoc. 100°	t	p dissoc. (3+1)
14.29	867(3+1)	-	-
16.39	865	-	-
19.23	867	69.0 (3+1)	214
23.81	868	78.5	347
25.32	472(3+2)	90.5	601
29.41	360	100.0	865
35.71	352	-	-
40.00	353	83 (3+2)	200
57.14	352	100	352
66.67	100	110	456
80.00	62	120	620

Ammonia (NH_3) + Cuprous iodide (CuI)

Lloyd, 1908

mol%	p dissoc.	mol%	p dissoc.
	51.75°		
13.74	771(3+1)	25.32	767
14.07	768	27.78	491(3+2)
14.71	766	31.25	489
15.61	767	35.71	490
16.67	768	43.48	487
17.54	765	55.56	489
18.87	767	66.67	490
20.83	769	74.07	180
23.80	768	83.33	90

t	p dissoc.	t	p dissoc.
	(3+1)		(3+2)
18.25	150	17.5	60
38.5	422	20.25	75
47.0	615	41.0	272
52.0	784	45.0	330
56.0	952	48.0	396
58.75	1084	51.75	490
		59.5	742
		65.0	999

Ammonia (NH_3) + Silver chloride (AgCl)

Horstmann, 1876

t	p dissoc. (3+1)	p dissoc. (3+2)
6	-	22.0
7	-	23.4
8	432	24.9
9	446	26.5
10	465	28.2
11	491	30.0
12	520	31.9
13	551	33.9
14	584	36.0
15	618	38.3
16	653	40.9
17	688	43.7
18	723	46.6
19	758	49.6
20	793	52.6
21	829	55.6

Jarry, 1897

t	p dissoc. (3+1)	p dissoc. (3+2)
0	262	12

Ammonia (NH_3) + Silver bromide (AgBr)

Jarry, 1898

t	p dissoc. (3+1) - (3+2)	t	p dissoc. (3+2) - AgBr
-23	140	0	880
-18	195	4	107
-8	360	10.4	166
0	605	11.2	174
3.5	745	14	206
8	920	16.6	249
14	1310	28.6	513
20	1820	43.8	1225
23	2140	53	1986

Ammonia (NH_3) + Silver iodide (AgI)

Fitzgerald, 1912

N	d	N	d
	-33.5°		
7.663	2.301	0.4847	0.7865
4.195	1.5740	.2832	.7690
2.009	1.114	.2637	.7385
1.850	1.0780	.2162	.7277
1.093	0.9185	.1036	.6970
1.012	0.8995	.06197	.6940

N	d	η	N	d	η
7.262	2.2275	9881	0.2016	0.7215	281.0
3.383	1.4250	920.8	.1136	.6900	268.7
1.984	1.1175	561.5	.0640	.6840	266.5
0.8460	0.8588	351.1	.03604	.6820	265.7
.3588	.7590	297.8			

N	d	N	d	N	d
1.849	1.0780	7.663	2.301	0.3951	0.7690
1.012	0.8145	4.1946	1.5740	.2378	.7277
0.4847	.7865	2.0088	1.114	.1036	.6970
.2637	.7385	1.0927	0.9185	.05634	.6940

Linhard, 1936

N	p	N	p
	0°		
0	3220.9	7.748	2999
0.274	3210	10.04	2903
0.542	3202	15.68	2602
0.921	3192	19.43	2373
2.066	3164	21.49	2242
3.313	3135	22.61	2166
4.646	3101	22.61	2165

Ammonia (NH_3) + Silver nitrate ($AgNO_3$)

Fitzgerald, 1912

N	d
-33.5°	
0.7375	0.7991
.4038	.7474
.1932	.7137
.1051	.6994

N	d	η
	-33.5°	
0.5921	0.7750	352.8
0.2759	0.7275	303.4
0.1618	0.7070	286.4
0.09788	0.6940	275.2
0.05724	0.6880	271.5
0.03264	0.6860	269.3
0.01915	0.6835	267.0

Juisa and Diamant, 1926

t	p dissoc.
(3+1)	
13.40	60.85
20.10	87.70
30.00	150.20
40.00	259.30
63,00	760.00
70.00	1001.10
80.00	1441.00

Ammonia (NH_3) + Magnesium chloride ($MgCl_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(4+1)		
-9	50	21	412
-3	98	23.5	472
0	122	27	585
4	140	29.3	660
7	165	32	748
11	210		
14.5	257		
19	352		

Ammonia (NH_3.) + Magnesium bromide ($MgBr_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
-20	155	-2	440
-17	181	-1	454
-14	229	0	480
-12.2	255	3	540
-11	275	5.8	600
-9	297	7.5	636
-8	307	9	675
-6	355	11	715
-5	380	14.5	795
-4	397		

Ammonia (NH_3) + Magnesium iodide (MgI_2)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
-18	86	14.5	400
-11	123	17	440
-8.5	144	20.8	510
-7	159	24	562
-4	180	27	620
0	215	31	697
5	260	32.3	719
8	305	34.5	760
11.5	350		

Ammonia (NH_3) + Calcium (Ca)

Kraus, 1908

t	p dissoc.	t	p	
+10.8	46.0 (6+1)	-32.5	471.8	sat.sol.
0	22.8	-50	192.8	
21.7	90.7			
+43.7	306.7			

Marshall and Hunt, 1956

mol%	p	mol%	p
-63.8°		-45.3°	
3.49	127.55	4.98	388.25
3.85	127.00	5.95	352.15
4.18	125.40	6.90	330.05
5.03	121.15	7.63	314.05
5.26	119.45	8.40	289.80
5.68	116.00	9.52	259.65
6.13	112.70	11.10	251.50
6.49	108.45	12.20	225.15
7.04	106.05	13.30	110.25
7.63	104.50	14.00	23.65
8.20	96.30	14.40	7.40
9.09	91.65	14.60	4.30
10.0	88.90	14.60	3.45
11.3	88.15	14.80	1.50
13.1	74.40	31.70	2.95
13.8	20.15		
14.5	3.25		
14.7	2.00		
14.9	0.75		
26.5	0.45		

t	p	t	p	t	p
$V+C_1+C_2$		$V+C_2$+sat.sol.		$V+L_1+L_2$	
-63.8	0.30	-63.8	74.40	-63.8	106.05
-45.3	1.80	-45.3	251.50	-45.3	330.00
-22.9	12.45				
0.0	45.95				

Birch and Mc Donald, 1948 (fig.)

mol%	f.t.
8.5	-78
10	-82
11	-85
12	-86

Ammonia (NH_3) + Calcium chloride ($CaCl_2$)

Ephraim, 1912

t	p dissoc.(8+1)	t	p dissoc.(8+1)
-5.5	71	20	380
0	93	24	483
6	156	28	620
10	201	33	800

Ammonia (NH_3) + Calcium nitrate (CaN_2O_6)

Kameyama, 1929

t	p sat.sol.	t	p sat.sol.
-40	306	-20	890
-30	517	-15	1128
-25	673	-10	1434

Linhard, 1936

molality	p	molality	p
	0°		
0	3220.9	2.962	3060
0.143	3214	3.633	2928
0.405	3207	4.335	2698
0.604	3203	4.780	2526
1.480	3183	5.01	2397
2.254	3144		

Portnov and Wassiliev, 1934

%	f.t.	%	f.t.
37.1(4+1)	-69.5	45.33	-9.5
37.39	-67	49.64	+14.0
38.54	-63	49.12	18.0 anh.
41.35	-51	51.98	48.5
43.17	-41	52.95	53.0
44.12	-19.5	58.12	70.0

Ammonia (NH_3) + Strontium (Sr)

Marshall and Hunt, 1956

mol%	p	mol%	p
	-63.8°		
1.86	121.65	16.1	9.35
2.11	120.45	16.47	1.90
2.70	118.16	16.50	1.95
2.90	117.85	16.55	2.05
3.71	115.10	16.61	2.95
5.46	114.40	16.80	1.50
6.49	115.35	16.89	0.65
10.6	92.35	16.92	.60
10.7	93.20	17.00	.95
13.3	93.40	17.10	.55
13.6	68.80	17.30	"
14.9	45.10	17.40	"
15.4	38.10	17.80	"
15.8	10.05	18.10	"
15.9	6.65	21.50	"
15.9	7.85	42.60	"
16.1	5.25		
16.1	4.40		
	-45.3°		
16.2	19.00	18.0	9.40
16.3	34.60	18.1	1.75
17.0	4.75	21.1	4.25
17.2	1.75	21.3	9.50
17.5	1.75	44.8	4.30

t	p
$V+C+C_2$	
-63.8	0.50
-45.3	3.10
-22.9	10.05
$V+C_2$+sat.sol.	
-63.8	93.30
$V+L_1+L_2$	
-63.8	115.0

Birch and Mc Donald, 1948

mol%	f.t.
5.5	-78
6	-81
6.5	-83
7	-89

Ammonia (NH_3) + Strontium nitrate (SrN_2O_6)

Linhard, 1936

molality	p	molality	p
	0°		
0	3220.9	1.530	3178
0.367	3207	1.736	3170
0.664	3201	1.91	3162
0.939	3195		

Ammonia (NH_3) + Barium (Ba)

Marshall and Hunt, 1956

mol%	p	mol%	p
	-63.8°		
2.44	127.65	10.2	55.75
2.63	127.30	10.4	53.80
2.90	125.55	10.6	45.60
3.20	125.35	10.8	36.00
4.15	124.70	10.8	34.00
4.83	120.60	11.0	21.80
5.49	116.50	11.2	13.40
5.92	111.95	11.2	8.90
6.25	108.75	11.3	6.65
6.90	101.50	11.5	2.10
7.58	93.05	11.7	2.35
8.26	87.70	11.7	0.70
8.62	87.75	11.7	0.65
9.26	87.70	11.8	0.90
9.43	84.50	12.0	0.20
9.52	74.90	12.1	0.20
10.0	55.70	12.2	0.20
	-45.3°		
4.12	402.15	10.4	155.65
5.43	358.75	10.7	101.15
7.63	279.45	11.1	49.40
8.55	252.25	11.4	10.00
	249.95	11.7	19.60
9.26	247.75	11.7	2.60
9.71	211.70	11.7	4.10
9.90	178.55	12.6	2.65
	173.30	18.9	2.95
10.2	168.10		

t	p
$V+C+C_2$	
-63.8	0.30
-45.3	2.65
-22.9	23.85
V+sat.sol.+C	
-63.8	55.75
-45.3	173.30
$V+L_1+L_2$	
-63.8	87.75
-45.3	249.95

Birch and Mc Donald, 1948

mol%	f.t.
(4+1)	
5.5	-78
7	-82
7.5	-87
7.7	-89

Ammonia (NH_3) + Barium chloride ($BaCl_2$)

Joannis, 1891

t	p dissoc.
0	541 (4+1)
28.4	1850

Ammonia (NH_3) + Barium nitrate (BaN_2O_6)

Linhard, 1936

molality	p
0°	
0	3220.9
0.288	3208
0.566	3201
0.833	3193

Portnov and Wassiliev, 1934

mol%	f.t.	mol%	f.t.
4.872(4+1)	-27	26.60	+2.5 anh.
10.34	-15.5	33.83	+8.7
12.07	-9	40.2	+12.7
14.4	-7	51.41	+22
21.28	-0.5	58.07	+51
22.31	+0.5		

Ammonia (NH_3) + Lead chloride ($PbCl_2$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(2+1)			
30	70 I	51	262 II
51	185	60	402
59.5	370	20	20 III
68.5	584	71.5	620
72.5	715	80	780
		72	630 IV

Ammonia (NH_3) + Lead bromide ($PbBr_2$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(2+1)			
-1	45 I	21	6 II
+22	134	43	94
32	201	52	180
42	299	61	470
52	430	66	548
53.5	455	70.5	675
60	561	74	753
62	600		
66.5	682	95	110
70	756	108	195
		123	254
		136.5	430
		140	450

Ammonia (NH_3) + Lead iodide (PbI_2)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(4+1)			
-15	40 I	0	118 II
- 3	120	7	200
+13	310	11	250
18.5	415	19	404
22	495	21	578
27	627	30	631.5
30.5	713	33	720
(2+1)			
42.5	125	70	500
47	160	78	659
56	275	81	745
57.5	290		
(1+1)			
83.5	64	126.5	460
100.5	140	135.5	611
115.5	287		
108.5	219	144	560
122.5	362	155.5	645
132.5	475		
156	405	155	200
156	235	154.5	184

Ammonia (NH_3) + Zinc chloride ($ZnCl_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
-13	6	50	479
0	22	53	545
10	45	56	636
19	82	57.5	682
27	131	60	767
39	268	61.3	801
45	367		

Ammonia (NH_3) + Zinc bromide ($ZnBr_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
0	11	45	235
10	26	46.3	255
13.5	34	48	278
23	62	52	348
25	70	56	425
31	100	60	510
32	107	64	627
39	165	67	716
40	172		

Ammonia (NH_3) + Zinc iodide (ZnI_2)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
43.5	268	60.5	636
49.5	367	61.5	682
55	479	63.5	767
57.5	545		

t	p dissoc.	t	p dissoc.
(6+1) second series			
11	37	38.5	196
14	45	44	275
17	54	48	342
21	66	49	356
28.5	106	54.2	468
34	148	58	563
38	194	63	750

Ammonia (NH_3) + Zinc chlorate ($ZnCl_2O_6$)

Ephraim and Johnsen, 1915

t	p	t	p
(4+1)		(6+1)	
114	16	-15	225
140	32	-14	232
165	50	-10	270
177	88	-8	304
		-6	332
		-2	388
		0	427
		+1	449
		3	486
		6	550
		7.5	594
		10	690
		11	747

Ammonia (NH_3) + Zinc iodate (ZnI_2O_6)

Ephraim and Johnsen, 1915

t	p
(4+1)	
77	44
91.5	81
106.5	141
119	242
130	356

Ammonia (NH_3) + Zinc nitrate (ZnN_2O_6)

Donskaya and Portnov, 1939

%	f.t.		T.C.D.
1.05	-79		+31
2.60	-79		+74
6.20	-79		+64
10.49	-78.5		+61
14.05	-78		-
16.75	-77		+61
18.14	-75	(10+1)	-
19.42	-73		+61
20.50	-72		-
20.86	-71		-
22.48	-67		-
24.18	-62		+63
25.16	-60		-
25.99	-58	(10+1)+(8+1)	-
25.16	-39	(8+1)	-
24.18	-31		-
22.48	0	(8+1)+(6+1)	-
24.18	+43	(6+1)	+63
25.16	48		-
26.03	58	(6+1)+(4+1)	-
27.44	60	(4+1)	-
28.77	63		+69
30.87	68		+71
33.46	75		+79
33.63	75.5		-
35.52	79		-
41.65	101		-
46.07	125		-

Ammonia (NH_3) + Zinc sulfate ($ZnSO_4$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(5+1)		(4+1)	
-19	33	82.5	166
-11	65	90.5	300
-7	88	95.5	387
0	126	100.5	500
+4	156	103.5	550
7.5	188	108.5	707
10	215	110	755
14	265		
15	275	91	261 after a time
19	351	97	326
22	400	104	440
25	472	113	640
27	540	117	725
30	615		
33	700	(3+1)	
36	790		
		102.5	175
		113	256
		121.5	290
		135	320

Ammonia (NH_3) + Cadmium chloride ($CdCl_2$)

Lang and Rigaut, 1899

t	p dissoc.	t	p dissoc.
(6+1)-(2+1)			
0	46	50	455
13	68	52	511
22	133	59	631
25	152	60	696
31	181	61	711
39	235	63	776
44	290	65	831
48	411	69	931

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
13.5	52	48.5	453
25	108	52	530
31	170	53.5	576
37.5	248	56	635
42	328	58	710
45	380	59.5	765

Ammonia (NH_3) + Cadmium bromide ($CdBr_2$)

Ephraim , 1912

t	p dissoc.	t	p dissoc.
(6+1)			
31	39	67	344
45	103	71	425
50	138	76	535
56	192	82.3	700
60.5	246	84.3	780

Ammonia (NH_3) + Cadmium iodide (CdI_2)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
18	3	91	369
24	8	94.5	435
44	30	99	525
59	60	99.5	542
64.5	96	103.5	633
69.5	126	104.5	666
78.5	204	107.5	733
85	271		

Ephraim, 1915

t	p dissoc.	t	p dissoc.
(6+1) second series			
66	103	86.5	344
71	138	99.3	535
77.5	192	106	700
83	246	108.5	780

Ammonia (NH_3) + Cadmium chlorate ($CdCl_2O_6$)

Ephraim and Johnsen, 1915

t	p	t	p
(6+1)		(4+1)	
69	89	116.5	117
82	156	125.5	174
94	241	136	335
103.5	330		
116.5	530		
122	772		

Ammonia (NH_3) + Cadmium iodate (CdI_2O_6)

Ephraim and Johnsen, 1915

t	p dissoc.
(4+1)	
91	381
101	486
110	709

Ammonia (NH_3) + Cadmium nitrate (CdN_2O_6)

Donskaya and Portnov, 1939

%	f.t.	%	f.t.
0.26	+37	14.18	-56
1.00	+4	18.88	-66
1.71	-9	19.62	-68
2.61	-21	22.50	-75
3.38	-25	22.78	-77.5
4.38	-29	10.32	-79.5
5.60	-33	6.01	-81.0
11.11	-47		

Ammonia (NH_3) + Cadmium sulfate ($CdSO_4$)

Isambert, 1870

t	p dissoc.
(3+1)	
48.5	368
51.5	439
100	1364

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(6+1)		(4+1)	
18	56 I	65	245
33	156	71.5	340
42	298	79	525
52	547	83	635
58.5	745	86.5	740
36	175	52.5	115 second series
47.5	341	61	200
54	436	73	385
		82.5	605

Ammonia (NH_3) + Mercuric iodide (HgI_2)

François, 1929

t	p dissoc.	t	p dissoc.
(2+1)		(4+3)	
0	8	15	1
15	18	25	2
25	37	35	3
35	72	45	6
45	130	55	12
55	219	65	23
65	362	75	39
75	601	85	65
80	732	95	107

Ammonia (NH_3) + Mercuric cyanide (HgC_2N_2)

Peters, 1912

t	p dissoc. 1	p dissoc. 2	t	p dissoc. 1	p dissoc. 2
		(2+1)			
59	88	-	25	15.5	14.5
55	73	76	20	8.5	11
50	57	58	19	6	8
45	45	45.5	15	6	5
40	34.5	34.5	10	4	5
35	25	27.5	5	3	-
30	18.5	20			

Brinkley, 1922

%	p	
0°		
369	88.25	(2+1)
556	88.37	
753	88.44	
993	88.41	
1053	83.43	(2+1)
1071	83.21	
1067	81.65	
1074	81.06	
1065	80.70	
1066	79.79	
1067	79.28	L + V
1098	79.07	
1200	78.13	
1257	77.64	
1362	76.60	
1445	75.67	
1536	74.80	
1615	73.90	

Franklin and Kraus, 1905

M	$\varkappa$ (in mhos)	M	$\varkappa$ (in mhos)
	-33°		
1.16	2.48	7.10	1.69
1.78	2.37	10.7	1.66
2.32	2.22	17.4	1.65
2.89	2.26	21.8	1.63
3.37	1.86	33.0	1.64
5.71	1.79	55.6	1.75

Ammonia (NH_3) + Potassium iodide Mercuric iodide (K_2HgI_4)

Peters, 1912

t	p dissoc.	t	p dissoc.
	(2+1)		
55	272	25	62
50	231.5	20	47.5
45	187	15	35
40	147.5	10	24
35	111	5	17.5
30	85	0	17

Ammonia (NH_3) + Cupric chloride ($CuCl_2$)

Ephraim, 1912-13

t	p dissoc.	t	p dissoc.
(6+1)		(4+1)	
63.5	114	83.5	65
71.5	180	96.5	139
78.5	252	106.5	239
85	323	116	375
88.5	380	125.5	629
95.5	508		
100	611	92	126 after a time
104.5	733	103	240
		108.5	340
		122	755

Ammonia (NH_3) + Cupric bromide ($CuBr_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
23	5	104.5	470
58	55	109.8	582
72.5	115	114.5	680
84	197	119	788
96.5	320		

Ammonia (NH_3) + Cupric iodide (CuI_2)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
54	100	106.5	450
70	188	112.5	545
82.5	264	119	707
93.5	342	120	740
102.5	408	120.5	754

Ammonia (NH_3) + Cupric chlorate ($CuCl_2O_6$)

Ephraim and Johnsen, 1915

t	p dissoc.	t	p dissoc.
(4+1)		(6+1)	
97	36	-15	201
116	55	-1	286
		+10	380
		20	491
		31	622

Ammonia (NH_3) + Cupric iodate (CuI_2O_6)

Ephraim and Johnsen, 1915

t	p dissoc.
(5+1)	
50.5	141
65	260
75	440
82	640
84.5	719

Ammonia (NH_3) + Cupric nitrate (CuN_2O_6)

Fitzgerald, 1912

N	d
-33.5°	
1.278	0.9085
0.699	.8069
.334	.7420
.182	.7157

N	d	η
1.326	0.9120	786.2
0.617	.742	424.6
0.266	.747	347.7

Ammonia (NH_3) + Cupric sulfate ($CuSO_4$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(5+1)		(4+1)-(2+1)	
20	3	120	62
70	123	137	123
87.5	375	151.5	320
95	578	160.5	502
100	721	166.5	665
		168.5	720

Ammonia (NH_3) + Manganese chloride ($MnCl_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
59	170	84	557
65	230	87.5	665
70	300	90.5	750
75.5	375	91	765
80	470		

Ammonia (NH_3) + Manganese bromide ($MnBr_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
83	86	120	468
90	127	124	550
104	240	128	650
110	308	130.5	720
114.5	370		

Ammonia (NH_3) + Manganese iodide (MnI_2)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
125.5	105	156	364
135	150	162.5	475
143.5	225	168.5	596
149.5	280	173.5	720

Ammonia (NH_3) + Manganese sulfate ($MnSO_4$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(6+1)		(2+1)	
40	86	102	73
53	166	114	99
67	450	128.5	130
73	700		
after a time			
50.5	261		
56.5	342		
65	476		
71.5	695		

Ammonia (NH_3) + Chromous chloride ($CrCl_2$)

Ephraim and Millmann, 1917

t	p dissoc.	t	p dissoc.
(6+1)		(3+1)	
0	104	44	45
13	194	59.5	97
16	228	72	198
27	387	86	433
33.5	517	92	618
43.5	744	95	730

Ammonia (NH_3) + Ferrous chloride ($FeCl_2$)

Huttig and Kollmann, 1951 (fig.)

mol%	P
208°	
33.3	0-10
25.0	15
20.0	20
17.7	20
14.3	20-158
12.5	158-162
11.1	162

t	P(dissoc.)	P(dissoc.)
	(8+1)-(7+1)	(7+1)-(6+1)
195	194	189
204	169	161
208	165	157
219	160	156
244	158	154
275	169	162

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
64.5	82	105.7	490
75	130	108.8	547
85.5	203	112.5	645
93	275	115	715
100.5	375		

Ammonia (NH_3) + Ferrous bromide ($FeBr_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
63	22	123.5	236
75	34	130.5	329
84.5	48	143	543
96.5	76	149	683
105.5	114	153	775
115.3	175		

Biltz and Fetkenhener, 1914

t	p dissoc.	t	p dissoc.
(2+1)			
136.5	6.9	183.5	42.5
137.0	7.1	230.0	82.4
153.5	13.3	234	85.2
168.0	27.0	236	89.9
(1+1)			
208	2.36	230	8.05
215	3.76		

Ammonia (NH_3) + Ferrous iodide (FeI_2)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
112	62	172.5	477
125	88	177.5	567
135.5	115	180.5	625
145.5	130	184	708
156.5	276	187.5	800
164	355		

Ammonia (NH_3) + Ferrous sulfate ($FeSO_4$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(6+1)			
73	150	98.5	551
82	256	102.5	642
88.5	350	107	740
94.5	464	109.5	788
after a time			
96	265	111.5	619
106	489	114.5	775
(4+1)-(2+1)			
100.5	182	128	492
113.5	368	140.5	565
118.5	422	159.5	639

Ammonia (NH_3) + Cobalt chloride ($CoCl_2$)

Peters, 1912

t	p dissoc.	t	p dissoc.
(6+1)-(2+1)			
50	123.5	25	65
45	115	20	55
40	102.5	15	46
35	87	16	34
30	75.5		

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
82.5	75	126	425
92.5	101	132	538
101.5	156	136	620
111	231	139.5	703
118	305		

Ammonia (NH_3) + Cobalt bromide ($CoBr_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
104	74	159.5	405
113	89	165	493
121.5	110	170.5	590
131.5	140	175	684
142	207	176.5	728
150	279		

Peters, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
50	52.5	25	15
45	45	20	12
40	37	18.5	13
35	29.5	15	10
30	22	10	8.5

Biltz and Fetkenheuer, 1914

t	p dissoc.	t	p dissoc.
(2+1)		(1+1)	
152.4	0.43	161.4	0.29
160.5	0.695	181.8	0.99
181.8	2.044	188.6	1.22
		208.4	3.46

Ammonia (NH_3) + Cobalt iodide (CoI_2)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
143.5	92	184.5	440
163	182	188.5	515
170	245	193	605
174.5	292	197.5	730

Peters, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
50	203.5	25	76
45	172	20	63.5
40	143	15	54.5
35	119.5	10	47
30	87	5	39

Biltz and Fetkenheuer, 1914

t	p dissoc.
(2+1)	
136.5	0.17
153.5	0.94

Ammonia (NH_3) + Cobalt sulfate ($CoSO_4$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(6+1)		(4+1)-(2+1)	
66	87	113.5	210 I
80.5	174	124	404
90	292	135	656
98.5	414		
102.5	484	103.5	141 II
106	538	112.5	256
109.5	596	124	443
		131.5	650
after a time		136	821
76	167		
87	264		
95.5	357		
105.3	495		
113.5	670		
117	770		

Ammonia (NH_3) + Nickel chloride ($NiCl_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
112	42	155	348
120	68	157.5	382
128	91	164	485
138	160	169	590
144.5	218	172.5	660
153	315	177	775

Peters, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
50	30	20	5.5
45	24	19	4.5
40	18	15	4.5
35	13.5	10	4
30	10	8	4
25	7.5		

Ammonia (NH_3) + Nickel bromide ($NiBr_2$)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
140.5	61	188.5	387
148.5	79	194	472
160.5	126	194.8	484
163.5	144	199	560
169.5	185	206.5	694
175	225	207.5	716
183.5	310	210	785

Peters, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
55	30	25	12
52	27	20	10
50	24	15	8.5
45	19	14	8
40	18.5	10	7.5
35	17	5	6.5
30	15		

Biltz, 1913

t	p dissoc.	t	p dissoc.
(6+1)		(2+1)	
84.5	4	195	2
92.5	5	207	3
104	6	218	8
117	9	236	14
134.5	26	250	26
150	45	262	37
152	56	273	68
158.5	73	278	81
163	101.5	283	95
170.5	147	289	102
174	161.5	295	142.5
180.5	203	299	146.5
192.5	352	304	168.5
198	487	309	191
200	531.5	310	199
203	598	311.5	266
206	708	318	284
213	828	325	321.5
		329.5	356.5
		334	441
		340	537
		347	598
		356	709

Biltz, 1914

t	p dissoc.	t	p dissoc.
(2+1)		(1+1)	
276.2	94.4	276.2	59.7
		299	121
		336	381

Ammonia (NH_3) + Nickel iodide (NiI_2)

Ephraim, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
170	84	213	362
182	125	220	470
195	195	225.5	565
205	275	234	720

Peters, 1912

t	p dissoc.	t	p dissoc.
	(6+1)		
50	37	25	16
45	32	20	11.5
40	28	15	8.5
35	24	5	5.5
30	20		

Ammonia (NH_3) + Nickel thiocyanate ($NiC_2N_2S_2$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
	(6+1)		
45	100	73.5	460
57	200	78	578
67	335	84	740

Next Page

Ammonia (NH_3) + Nickel nitrate (NiN_2O_6)

Ephraim , 1913

t	p dissoc.	t	p dissoc.
(6+1)			
115	66	122.5	90 after a time
135.5	124	140	165
147.5	193	153.5	234
159.5	278	168.5	379
171.5	412	177	492
180	512	185.5	633
		191	757

Ephraim , 1913

t	p dissoc.	t	p dissoc.
(5+1)			
47.5	60	88.5	333
62	125	99.5	464
69	165	106.5	575
75	208	115	724
85.5	295		

Ammonia (NH_3) + Nickel chlorate ($NiCl_2O_6$)

Ephraim, 1913 and Ephraim and Johnsen, 1915

t	p dissoc.
(6+1)	
126	50
140.5	75
159	140

Ammonia (NH_3) + Nickel iodate (NiI_2O_6)

Ephraim and Johnsen, 1915

t	p dissoc.
(5+1)	
53	179
65	365
77	670
79	905

Ammonia (NH_3) + Nickel sulfate ($NiSO_4$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(6+1)		(4+1)-(2+1)	
100.5	170	136	88
111	257	151.5	143
122.5	441	163.5	350
131.5	673	170.5	550
134	752	175	717
after a time			
91	62		
99	110		
109.8	190		
120	294		
123.5	338		
130	442		
135	525		
141.3	680		
144.5	810		

Peters, 1912

t	p dissoc.	t	p dissoc.
(6+1)			
50	82	25	24
45	67.5	20	17.5
40	53	15	14
35	39.5	10	12.5
30	31.5	5	11

Ammonia (NH_3) + Nickel thiosulfate (NiS_2O_3)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(5+1)			
91.5	80	132.5	510
109.5	168	137	600
121.5	312	142	715

Previous Page

Ammonia (NH_3) + Nickel dithionate (NiS_2O_6)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(6+1)			
116	70	169	430
130.5	142	178.5	556
148	245	182.5	646
160	340		

Ammonia (NH_3) + Nickel tetrathionate (NiS_4O_6)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(6+1)			
76	60	126	590
96.5	158	130	675
105.5	253	131.5	699
116.5	440		

Ammonia (NH_3) + Nickel hypophosphite ($NiH_4O_4P_2$)

Ephraim, 1913

t	p dissoc.
(6+1)	
40	40
51	95
57.5	146
62	185

Ammonia (NH_3) + Nickel bicarbonate ($NiC_2H_2O_6$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(6+1)		(4+1)	
0	125	60	9
10	218	102.5	60
16	306	118.5	146
19	355	131.5	375
20	364	140	775
24	450		
31	620		
35	745		

Ammonia (NH_3) + Nickel acetate ($NiC_4H_6O_4$)

Ephraim, 1913

t	p dissoc.	t	p dissoc.
(4+1)			
20	20	73.5	440
31	60	77	522
43	125	79.5	610
53.5	186	80	637
63	265	84	780
66.5	325		

Ammonia (NH_3) + Platinum chloride ($PtCl_2$)

Ephraim and Millmann, 1917

t	p dissoc.
(5+1)	
-14	100
+ 4	300
+13	500
+22	760

Ammonia (NH_3) + Platinum iodide (PtI_2)

Ephraim and Millmann, 1917

t	p dissoc.	t	p dissoc.
(6+1)		(4+1)	
-8	565	126	130
-7	585	163	588
-5.5	668		

Ammonia (NH_3) + Aluminum chloride ($AlCl_3$)

Baud, 1901

t	p dissoc.
(9+1)	
0	1790
-10.7	970
-22.3	481
-37	189

Ephraim and Millmann, 1917

t	p dissoc.
(6+1)	
47	19
69.5	39
97.5	140
123.3	518
128	690

Ammonia (NH_3) + Aluminum bromide ($AlBr_3$)

Klemm and Tanke, 1931

mol%	f.t.	mol%	f.t.
0	100	45	115
5	98	50	125 (1+1)
10	95	55	117
20	82	56	100
30	90	60	170
40	100	65	270

Ephraim and Millmann, 1917

t	p dissoc.
(6+1)	
148	86
171.5	118
196.5	244
211	440
229	788

Ammonia (NH_3) + Aluminum iodide (AlI_3)

Ephraim and Millmann, 1917

t	p dissoc.
(6+1)	
19.5	30
52.5	161
80.5	801

Ammonia (NH_3) + Chromic bromide ($CrBr_3$)

Ephraim and Millmann, 1917

t	p dissoc.
(6+1)	
72.5	20
110	84
142	411

Ammonia (NH_3) + Chromic sulfate ($Cr_2S_3O_{12}$)

Ephraim and Millmann, 1917

t	p dissoc.
(6+1)	
70	60
110	115
144	283
158	493
171.5	758

Ammonia (NH_3) + Chromic nitrate dibromide ($CrNO_3Br_2$)

Ephraim and Millmann, 1917

t	p dissoc.	t	p dissoc.
(6+1)			
138.7	77	176	458
161	218	186	743

Ammonia (NH_3) + Ferric chloride ($FeCl_3$)

Ephraim and Millmann, 1917

t	p dissoc.	t	p dissoc.
(6+1)			
27.5	112	58	468
38.3	200	65	586
49	324	70.5	714

Ammonia (NH_3) + Ferric bromide ($FeBr_3$)

Ephraim and Millmann, 1917

t	p dissoc.	t	p dissoc.
(6+1)			
0	73	40.2	490
16	156	49	661
30.5	331	53	734

Ammonia (NH_3) + Ferric sulfate ($Fe_2S_3O_{12}$)

Ephraim and Millmann, 1917

t	p dissoc.	t	p dissoc.
(12+1)			
0	80	29.3	475
8.5	127	36	705
16.2	194	37	745
22.2	290		

Ammonia (NH_3) + Potassium Ferric chloride (K_2FeCl_5)

Ephraim and Millmann, 1917

t	p dissoc.	t	p dissoc.
(1+1)			
0	130	26.5	450
14	275	32	560
20.5	365	40.5	720

Ammonia (NH_3) + Cerium chloride ($CeCl_3$)

Barre, 1913

t	p dissoc.	t	p dissoc.
(20+1)		(4+1)	
-69	66	50	224
-40	558	60	490
-35	678	65	652
-26	960	70.5	919
-16.5	1524	75	1288
0	3148		
(12+1)		(2+1)	
-30	328	90	414
-18	635	100	828
-15	710	105	1211
-11	882		
0	1572		
(8+1)			
0	262		
9.5	497		
19	800		
27	1080		
30	1305		

Ammonia (NH_3) + Thallic chloride ($TlCl_3$)

Ephraim and Millmann, 1917

t	p dissoc.
(3+1)	
21.5	24
34	53
53	108
77.5	240
101	688

Ammonia (NH_3) + Thallic sulfate ($Tl_2S_3O_{12}$)

Ephraim and Millmann, 1917

t	p dissoc.
(1+1)	
60	173
75	310
100	770

Ammonium chloride (H_4NCl) + Lithium chloride (LiCl)

Hachmeister, 1919

mol%	wt%	b.t.	f.t.	E	min	tr.t.
0	0	338	-	-	-	174
8.86	7.17	332	-	264	50	166
24.68	20.61	329	-	266	140	170
32.55	27.65	331	-	267	180	168
38.10	32.77	334	320	264	200	"
40.48	35.00	335	313	267	220	"
41.11	35.50	"	310	265	230	166
45.69	40.00	340	291	"	240	168
49.61	43.82	342	-	267	300	-
52.45	46.62	"	280	270	270	-
62.18	56.58	350	342	"	"	-
100	100	-	605	-	-	-

Kordes, 1926

mol %	f.t.
100	613
-	267 E

Ammonium chloride (NH_4Cl) + Potassium chloride (KCl)

Rassow, 1920

wt%	mol%	f.t.	m.t.
0	0	520	520
6.83	5	556	552
13.41	10	570	561
19.74	15	576	571
25.84	20	590	576
31.73	25	610	592
36.85	30	643	602

Ujeda, 1913

% C	% C	t
13	38	90
10	62	65
7	79	25

Jänecke, 1928 and 1914

mol %	f.t.	mol %	f.t.
0	520	23	560
7	539	32	584
15	546	41	603

Krickmeyer, 1896

%	d	%	d
	20°		
100	1.994	3.24	1.544
96.98	1.976	0	1.532

Ammonium chloride (NH_4Cl) + Cuprous chloride (CuCl)

Hachmeister, 1919

mol%	wt%	b.t.	f.t.	E	min	tr.t.
0	0	-	-	-	-	174
3.21	5.75	-	-	-	-	182
17.31	7.93	-	-	-	-	180
27.80	9.17	-	-	138	10	181
32.12	46.69	338	-	"	30	184
36.49	51.55	339	305	136	70	179
41.10	56.36	340	280	140	130	180
46.30	61.48	-	253	134	170	179
52.80	67.36	345	201	140	160	-
58.86	72.60	351	170	"	330	-
66.23	78.39	361	196	141	300	-
72.00	80.90	-	244	140	240	-
78.47	87.08	378	320	"	140	-
84.50	90.95	394	364	"	70	-
91.01	94.94	415	390	136	40	-
100	100	954	425	-	-	-

Kordes, 1926

mol %	f.t.
100	425
-	140 E

Ammonium chloride (NH_4Cl) + Silver chloride (AgCl)

Jänecke, 1923 (fig.)

%	f.t.	E 1	E 2	tr.t. 1	tr.t. 2	tr.t. 3
100	465	-	-	455	375	180
95	455	245	-	"	"	"
80	375	"	-	"	"	"
75	365	"	-	"	"	"
50	300	"	-	"	"	"
50	342	"	-	"	"	"
30	265	"	-	"	"	"
20	245	"	-	"	"	"
18	265	"	-	"	"	"
15	300	"	-	"	"	"
9	335	"	335	"	"	"
5	455	"	"	"	"	"
4	375	"	"	-	"	"
2	335	"	"	-	-	"
0	400	-	-	-	-	"

Ammonium chloride (NH_4Cl) + Mercurous chloride (HgCl)

Jänecke, 1923 (fig.)

%	f.t.	E 1	E 2	tr.t. 1	tr.t. 2
0	340	-	-	-	180
15	328	328	235	-	"
25	336	"	"	336	"
50	364	"	"	"	"
75	328	"	"	"	"
79	235	-	"	"	"
85	336	-	"	"	"
100	383	-	-	"	"

Ammonium chloride (NH_4Cl) + Mercuric chloride ($HgCl_2$)

Jänecke, 1908 and 1923

mol%	wt%	b.t.	f.t.	E	tr.t. 1	min	tr.t. 2	min
100	100	310.5	282.5	-	-	-	-	-
79	95	324	246	-	-	-	-	-
64	90	347	184	184	-	-	-	-
50	83.5	-	204(1+1)	-	-	-	-	-
44.2	80	350	200	199	-	-	-	-
33.3	71.7	-	236	198	-	-	-	-
31.6	70	344	244tr.t.	200	-	-	-	-
25	67.2	-	291(4+1)	200	248	-	-	-
22.8	60	338	-	195	242.5	7	-	-
20	55.9	-	330	-	247	-	-	-
16.4	50	338.5	-	-	243.5	5	180	3
12.2	40	338	-	-	246	3	177	4
7.8	30	336.5	-	-	235	3	170	3
4.5	20	337	-	-	242.5	2	180	3
2.2	10	339	-	-	239	1	180	9
0.0	0	340	-	-	-	-	180	11

Belyaev and Mironov, 1952

mol%	f.t.	mol%	f.t.	mol%	f.t.
100.0	283.0	67.0	214.5	48.0	215.0
97.5	280.0	66.5	214.0(1+2)	46.0	214.0
95.0	275.5	66.0	212.0	45.0	213.0
92.5	271.5	65.0	210.0	43.0	210.0
90.0	268.0	64.8	209.5	41.0	203.0
87.5	263.0	63.0	207.0	40.0	201.5
85.0	258.0	62.8	207.0	39.0	193.0 E
83.5	253.0	61.5	205.0	37.5	209.0
82.5	249.5	61.3	198 E	36.5	220.0
82.0	246.0	61.0	202.0	35.0	230.0
81.5	243.0	60.0	209.5	34.0	237.5
80.8	242.0	59.0	209.0	33.5	238.5
80.0	240.5	58.0	208.0	33.4	243.0(2+1)
79.0	238.5	57.5	209.5	33.0	242.0
77.5	237.0	56.0	212.5	32.0	251.0
76.5	235.0	55.0	213.5	31.0	262.0
72.5	235.0(1+3)	53.0	214.0	30.0	273.5
70.0	232.0	51.5	217.5	29.0	296.0
68.5	226.5	50.0	218.0(1+1)	27.0	316.0
67.5	219.5				

Belyaev and Mironov, 1952

mol%	κ (in mhos) 225°	250°	275°	300°
100.0	-	-	-	0.000124
94.9	-	-	0.090	.111
90.0	-	0.219	.231	.242
84.1	-	.284	.383	.491
83.0	-	.394	.486	.577
78.5	-	.537	.625	.711
78.0	-	.648	.718	.790
73.0	0.657	.779	.889	.985
72.0	.504	.650	.786	.891
70.0	.432	.561	.682	.787
66.6	.547	.688	.817	.940
62.4	.616	.790	.930	1.048
56.8	.602	.787	.970	.131
54.8	.618	.784	.940	.075
52.7	.458	.700	.923	.1116
50.3	.591	.802	.976	.111
50.0	.730	.785	1.080	.122
48.5	.707	.928	.077	.183
46.2	-	-	.046	.138
44.2	.702	.890	.040	.137
41.8	.707	.378	.023	.113
38.0	.678	.838	0.953	.023
35.0	.630	.790	.925	.021
32.2	-	.702	.855	.960
29.4	-	.877	.948	1.010

Ammonium chloride (NH_4Cl) + Zinc chloride ($ZnCl_2$)

Grünauer, 1904

%	f.t.	%	f.t.	%	f.t.
100	365	83.59	220	64.54	270
96.56	290	78.10	200	56.02	300
92.72	250	71.81	175	45.92	335
88.43	235				

Hachmeister, 1919

mol%	wt%	b.t.	f.t.	E	min	tr.t.
0	0	338	-	-	-	174
5.50	3.00	342	-	-	-	168
12.20	16.11	"	-	-	-	"
20.20	39.21	"	-	-	-	167
25.10	46.08	340	-	-	-	-
33.39	42.10	341	-	-	-	-
40.31	63.24	342	-	-	-	166
45.79	68.27	343	279	179	140	-
50.70	72.39	"	197	182	280	-
58.20	78.01	345	222	180	100	-
63.10	79.33	-	244	179	40	-
67.19	83.91	348	249	-	-	-
69.85	85.51	-	245	232	100	-
72.81	87.21	-	-	"	250	-
75.07	88.47	358	240	"	200	-
79.35	90.73	-	247	230	100	-
83.77	92.93	-	256	228	50	-
86.96	94.44	385	261	"	20	-
100.00	100.00	430	283	-	-	-

Ammonium chloride (NH_4Cl) + Cadmium chloride ($CdCl_2$)

Hachmeister, 1919

mol%	wt%	b.t.	f.t.	(1+2)	E_1	min	E_2	min
0	0	338	-	-	-	-	-	-
28.04	57.18	340	-	-	-	-	-	-
39.12	68.77	341	-	-	268	10	-	-
43.08	72.08	346	-	-	"	60	-	-
47.98	75.95	356	-	-	266	150	-	-
49.87	77.32	-	289	-	-	-	-	-
52.30	78.98	375	-	-	266	180	-	-
58.93	83.10	403	308	-	270	"	-	-
65.51	86.68	425	363	-	267	120	-	-
66.94	87.40	430	390	362	"	100	312	20
68.41	88.12	-	400	366	268	80	308	"
70.94	89.32	440	426	363	267	70	312	40
71.41	89.54	447	433	357	"	60	"	50
73.01	90.26	455	447	348	268	70	"	60
74.43	90.89	464	455	337	267	"	"	70
76.34	91.71	472	466	332	268	"	315	90
80.59	93.43	492	485	-	267	60	312	80
88.35	96.29	540	522	-	"	30	"	20
91.69	97.42	555	535	-	266	20	313	10
92.56	97.71	572	542	-	270	10	315	10
95.52	98.65	-	550	-	-	-	312	"
100.00	100.00	964	568	-	-	-	-	"

Ammonium chloride (NH_4Cl) + Antimony chloride ($SbCl_3$)

Kendall, Crittenden and Miller, 1923

%	f.t.	%	f.t.
97.9	71.3	65.3	134.8
97.4	70.8	63.8	139.7
95.6	69.7	62.8	157.4
93.4	68.1	57.8	217.8
92.5	67.5	56.8	228.6
92.7	68.2	52.6	253.5
92.5	68.0	44.7	284.3
91.8	66.5	41.1	290.5
91.1	66.6	38.3	289.5
90.1	65.7	36.5	289.5
87.5	64.4	35.4	296.8
85.9	63.1	33.3	305.8
81.9	59.5	31.3	318.0
81.2	70.0	27.7	338.5
77.1	89.2	18.7	392.0
74.5	104.1	14.1	418.0
72.4	110.6	0	73.4
68.7	123.1		
		(1+1)	(3+2)

Ammonium chloride (NH_4Cl) + Aluminum chloride ($AlCl_3$)

Kendall, Crittenden and Miller, 1923

%	f.t.	%	f.t.
99.80	192.2	56.5	274.1
79.4	192.3	50.9	301.2
76.2	186.6	46.9	287.2
71.5	171.8	43.2	266.1
68.1	157.9	41.6	257.1
66.8	173.4	39.0	284.3
63.1	229.4	34.9	357.0
61.8	239.7		

Ammonium chloride (NH_4Cl) + Ferric chloride ($FeCl_3$)

Hachmeister, 1919

mol%	wt%	b.t.	f.t.	E	min	tr.t.
0	0	338	-	-	-	174
9.57	24.29	342	-	234	60	157
14.23	33.49	"	-	236	50	158
29.84	56.33	340	-	235	90	-
34.74	61.74	342	310	235	120	-
36.26	63.28	"	263	232	150	-
40.76	67.61	345	244	233	160	-
42.50	69.15	-	252	232	120	-
45.08	70.50	347	261	235	90	-
48.56	74.11	-	296	-	-	-
49.26	74.65	370	297	-	-	-
51.97	76.60	386	295	-	-	-
60.29	82.16	-	253	221	150	-
63.30	83.95	338	238	225	200	-
69.11	87.15	-	255	221	180	-
73.64	89.44	-	263	221	120	-
74.55	89.88	-	265	220	120	-
76.28	90.70	-	267	219	100	-
79.97	92.37	-	275	221	90	-
84.93	94.47	-	282	220	50	-
100.00	100.00	307	303	-	-	-

(1+1)

Ammonium bromide (NH_4Br) + Mercuric bromide ($HgBr_2$)

Belyaev and Mironov, 1952

mol %	f.t.	tr.t. I	tr.t. II
100.0	243.0	-	-
95.9	239.5	-	-
95.0	238.0	222	182
91.6	237.0	-	-
90.0	235.0	222	198
85.0	232.0	222	202
82.5	227.0	222	200
81.9	226.0	-	-
80.0	225.0	222	155
77.5	221.5	-	155
76.5	221.5	-	-
75.0	221.0	-	155
72.5	220.0	-	-
70.0	216.0	-	155
68.7	212.0	-	-
67.5	208.0	-	155
66.6	205.5	-	-
65.4	200.5	-	-
65.0	199.0	-	155
62.5	189.0	-	155
61.0	170.0	-	-
60.0	159.0	-	155
57.5	-	-	140
56.8	153.5	-	-
55.0	151.0	-	140
53.8	148.0	-	-
52.5	140.0	-	140 E
52.1	146.0	-	-
50.1	165.0	-	-
50.0	166.0	-	140
48.0	177.5	-	-
46.4	184.0	-	-
45.0	189.0	-	140
44.8	190.5	-	-
42.7	194.0	-	-
42.5	-	-	140
40.8	199.0	-	-
40.0	200.0	-	-
39.0	208.0	-	-
37.3	227.0	-	201
35.9	238.0	-	-
35.0	248.0	-	201
32.5	-	266	201
30.0	316.5	267	201
25.0	371.0	268	-

Ammonium bromide (NH_4Br) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1928

mol%	f.t.	mol%	f.t.
99.48	160.5 L_1+L_2	67.6	103.8
99.20	236.5	67.1	104.0 (1+2)
79.60	159.5	66.5	103.6
76.00	94.8	63.7	143.8
77.20	96.6	61.5	166.1
76.90	96.9	57.0	197.2
76.20	97.5 (1+3)	53.9	214.0 (1+1)
74.80	97.3	50.5	230.5
72.70	92.7	49.3	229.8
72.70	98.1	46.6	213.6
71.80	99.9	45.3	207.5
69.00	103.2	42.9	360.0
		0	97.1

Gorenbein, 1947

%	d				
	110°	120°	130°	140°	150°
84.50	2.699	2.687	2.673	2.660	2.647
86.90	.711	.697	.683	.669	.655
88.98	.717	.703	.688	.674	.659
89.42	.716	.699	.685	.668	.656
90.81	.713	.697	.682	.665	.649
100.00	.624	.599	.578	.555	.532

%	η				
	110°	120°	130°	140°	150°
84.50	-	14483.5	11729.1	9755.8	8338.3
86.90	17095.8	13797.1	11122.5	9312.4	7877.3
88.98	16070.1	12905.8	10613.4	8898.0	7682.7
89.42	15310.3	12353.6	10090.2	8547.4	7185.5
90.81	13506.3	11032.0	9127.4	7724.0	6650.9
100.00	2167.9	1946.2	1762.1	1610.0	1453.1

%	κ				
	110°	120°	130°	140°	150°
84.50	220.030	270.220	325.36	372.820	426.450
86.90	191.718	228.939	271.78	317.772	364.090
88.98	168.143	198.734	229.89	273.960	312.811
89.42	158.071	188.666	217.05	254.738	286.149
90.80	139.036	168.393	197.28	225.074	249.747

Ammonium Selenium bromide ($N_2H_8Br_6Se$) + Ammonium Stannic bromide ($N_2H_8Br_6Sn$)

Carozzi, 1924

%	d
100.00	3.51
73.80	.50
51.24	.43
27.00	.39
0.00	.328

Ammonium selenium bromide ($N_2H_8Br_6Se$) + Ammonium platinum bromide ($N_2H_8Br_6Pt$)

Carozzi, 1924

%	d	%	d
100.00	4.200	35.17	3.624
79.96	3.905	16.80	.437
64.90	.861	0.00	.328
49.00	.705		

Ammonium iodide (NH_4I) + Mercuric iodide (HgI_2)

Belyaev and Mironov, 1952

mol%	f.t.	E	tr.t.
100.0	257.5	-	127
95.0	254.0	-	121
92.3	250.0	-	-
90.0	248.0	-	120
89.7	248.0	-	-
88.0	246.0	-	-
86.8	242.5	-	-
86.3	242.0	-	-
85.7	240.0	-	-
85.0	237.5	-	-
84.9	238.0	-	-
84.6	236.0	-	-
83.4	235.0	-	-
82.4	232.0	-	-
82.0	231.5	-	-
81.3	231.0	-	-
80.6	230.0	-	-
80.4	228.5	-	-
79.3	227.0	-	120
77.4	219.0	-	-
75.6	215.0	113	120
74.4	210.0	-	-
72.5	205.0	-	-
70.5	197.0	-	-
69.0	193.0	113	120
66.2	180.0	-	-
65.0	160.0	113	120
62.5	145.5	-	-
61.5	140.0	-	-
60.6	136.0	-	-
59.9	132.0	113	120
59.2	126.0	-	-
58.5	123.0	-	-
57.8	123.0	-	-
57.1	119.0	-	-
56.9	116.0	-	-
56.5	113.0	-	-
56.1	126.0	-	-
55.0	139.0	113	-
53.8	150.5	-	-
52.7	160.0	-	-
51.6	169.0	-	-
50.5	176.5	-	-
50.0	182.0	-	-
47.7	194.5	-	-
45.8	201.5	-	-
43.9	207.0	-	-
41.0	210.5	-	-
40.4	211.0	-	-
40.0	211.0	-	-
39.7	216.0	-	-
39.0	219.5	-	-
38.2	227.0	-	-
37.9	231.5	-	-
37.5	236.0	-	-
37.1	239.0	-	-
36.8	242.5	-	-
36.7	243.0	-	-
35.0	278.5	-	210
33.8	288.0	-	-
29.9	344.0	-	234
27.5	385.0	-	-
25.2	388.0 b.t.	-	-

Ammonium nitrate ($N_2H_4O_3$) + Lithium nitrate ($LiNO_3$)

Campbell, 1942

%	f.t.	tr.t.	E
100	270	-	-
75	218	-	-
70	212	-	-
65	205	-	-
60	196	116	-
55	182	118	95
50	171	120	97
45	151	119	96
40	135	117	96
35	121	-	97
30	110	-	"
25	97	-	"
20	106	-	"
15	119	-	"
10	133	126	"
5	153	128	93
0	169	126	-

Holmes,Jr., O'Connell and Hankard, 1951

%	f.t.	
	I	II
0	170	170
10	138	138
15	122	117
20	109	101
25.3	95	79.5 E unstable
26.6	90.4	90.4 E stable
30	114	114
35	140	140
40	153	153
50	174	174
60	194	194

Ammonium nitrate ($N_2H_4O_3$) + Sodium nitrate ($NaNO_3$)

Early and Lowry, 1922

%	f.t.	E	%	f.t.	E
0	169.6	-	22.0	124.8	120.8
5.0	156.5	-	23.0	126.6	120.6
10.0	144.2	-	23.5	129.1	120.7
15.0	132.0	120.6	25.0	133.4	120.8
17.5	126.3	"	30.0	147.0	120.2
18.5	124.3	120.8	40.0	174.2	119.2
20.0	121.9	"			

%	tr.t. I	tr.t. II	tr.t. III
0	125.2	84.2	32.2
5	125.4	83.6	-
10	125.5	83.3	31.6
20.5	-	82.0	31.4
25	-	82.5	-

Holmes Jr. and Revinson, 1944

%	f.t.	tr.t.	E
7.5	151.7	126.2	117.2
14.0	135.3	126.2	120.6
19.0	124.0	-	121.2
20.9	121.6	-	121.2
25.0	137.7	-	121.2
30.0	153.1	-	120.9

%	f.t.	%	f.t.	
4.5	158.6	18.0	126.7	tr.t.
7.5	151.7	19.0	124.4	
9.5	146.3	20.0	122.9	
10.0	145.7	21.0	121.9	E
10.5	145.0	22.1	125.0	
11.2	143.2	23.7	132.3	
12.5	139.9	25.4	138.3	
14.0	135.5	27.2	143.5	
15.0	133.0	30.0	153.2	
16.5	129.4	35.2	167.2	
17.0	128.2			

Dingemans and Djkgraaf, 1948

%	E I	E II
21.8	121.1	117.1

Ammonium nitrate ($N_2H_4O_3$) + Potassium nitrate (KNO_3)

Perman and Saunders, 1923

%	f.t.	%	f.t.
0.0	169.5	15.88	160.0
4.0	166.0	16.0	160.1
8.0	162.8	17.49	164.0
13.4	157.3	18.3	165.9
15.0	158.6	19.0	167.9

E : 13 6 % 156.5

Jänecke, 1928 (fig.)

mol%	f.t.	m.t.	tr.t.I begin	tr.t.I end	E_1
0	167	167	130	130	-
2	166	162	128	125	127
3	165	160	127	123	"
5	163	158.5	131	119	"
7	161	157	-	115	"
9	160	"	-	110	"
16	157	"	-	131	"
25	185	"	-	"	"
37	220	"	-	"	"
41	235	"	-	"	"
44	240	"	-	"	"
48	248	"	-	"	"
50	249.7	"	-	"	"
75	300	"	-	"	"
78	305	"	-	"	"
84	313	"	-	128	-
88	320	"	-	126	-
92	325	"	-	125	-
97.5	332	131	131	124	-
98	334	157	130	123.5	-
100	335	335	123	123	-

mol%	E_2	tr.t.II begin	tr.t.II end	tr.t.III begin	tr.t.III end	tr.t.IV begin	tr.t.IV end
0	-	85	85	35	35	-16	-16
2	-	93	"	25	0	-18	-20
3	-	97	"	16	-20	-20	-48
5	-	102	85.5	10	"	"	"
7	110	106	86	8	"	"	"
9	"	110	87	-4	"	"	"
16	"	-	89	-14	"	"	"
25	"	-	93	-36	"	-	"
37	"	-	101	-48	"	-	"
41	"	-	103	0	"	-	"
44	"	-	108	+50	-	-	-
48	"	-	110	110	-	-	-
50	"	-	-	-	-	-	-
75	"	-	-	-	-	-	-
78	"	-	-	131	-	-	-
84	"	-	-	0	-	-	-
88	"	-	-	-30	-	-	-
92	"	-	-	-48	-	-	-
97.5	"	-	-	-	-	-	-
98	-	-	-	-	-	-	-
100	-	-	-	-	-	-	-

Morand, 1955

%	IV	III (by heating)	II (by heating)	I (tr.t.)	I (by cooling)	II-IV (by cooling)
100	-18	32	82	125	125	55
98	-17	29	86	122	122	56
96	-15	22	90	122	122	59
95.8	-13	-	99	123	123	80
90	-	-	101	121	121	88
85.8	-	-	105	117	117	92
80	-	-	106	"	"	93
70	-	-	"	"	"	"

Whetstone, 1948

mol%	t	a	b	c
		Lattice constants (in Å)		
0 (III)	42	7.14	7.67	5.84
19	20	7.06	7.60	5.79

Ammonium nitrate ($N_2H_4O_3$) + Cesium nitrate (NO_3Cs)

Morand, 1955

%	III (by heating)	II (by heating)	I (tr.t.)	I (by cooling)	II b (by cooling)
100	32	82	125	125	55
97.25	"	80	"	"	45
96.65	"	67.5	"	"	37
96.40	"	40	"	"	32
95	"	38	123	123	"
89.20	"	"	113	113	"
83	-	-	105	105	-

Ammonium nitrate ($N_2H_4O_3$) + Thallium nitrate (NO_3Tl)

Boks, 1902

mol %	f.t.	m.t.
0	165.2	164.6
9.8	171.8	170
19.99	177.8	175.5
29.86	182.6	185
44.66	188.8	186.5
59.98	194.6	-
68.2	197.8	196
80	201	199.6
90	203.6	261.5
100	204.8	204.1

mol % at f.t. L	mol % at f.t. C
69	73
49.9	54.9
21.2	27.7

mol%	tr.t.I	mol%	tr.t.I
1st method		2nd method	
0	122.6	0	126.2
10	116.4	9.92	116.2
20.0	106.0	19.67	109.2
29.86	110.8	23.98	110.4
40.01	112.7	29.38	111.8
55.93	123.8	41.35	117.2
80	131.4	72.8	131.2
100	143.2	100	146.6

mol%	t.tr.(II)	mol%	t.tr.(II)
0	85.5	41.45	75.6
1.14	83	72.8	75.8
2.84	78.4-85	92.97	76.2
23.98	75-77	95.98	77.6
29.83	75.6	100	79.5

Morand, 1955

%	III (tr.t. by heating)	IIa	II	I
100	32	-	82	125
95.17	32	-	80	125
92.70	31	-	76	122
90.45	32	-	67	115
87.32	32	37.5	58	113
85	32	37.5	-	109
82.10	32	37.5	-	105
79	-	37.5	-	105

%	I (tr.t. by cooling)	IIb	IIa
100	125	55	-
95.17	125	51	-
92.70	122	49	-
90.45	115	45	-
87.32	113	36	37.5
85	109	32	"
82.10	105	32	"
79	105	-	"

Ammonium nitrate ($N_2H_4O_3$) + Silver nitrate ($AgNO_3$)

von Zawidzki, 1904

mol%	wt%	f.t.	mol%	wt%	f.t.
0.00	0.00	167.8	53.58	71.00	113.8
5.08	10.20	154.6	57.99	75.54	123.8
8.66	16.75	145.0	60.71	76.63	130.2
16.77	29.96	124.2	64.65	79.52	139.4
20.27	35.04	118.8	71.53	84.21	153.7
24.52	40.81	111.5	75.53	86.76	161.4
27.46	44.55	105.7	78.05	88.30	166.3
34.92	53.24	105.2	84.76	92.19	180.3
39.92	58.51	107.4	90.63	95.35	191.8
45.32	63.76	109.3	95.23	97.70	200.6
50.00	67.97	109.6	100.00	100.00	209.0
50.42	68.34	109.5			

E_1 : 101.5° E_2 : 109.6°

tr.t. NH_4NO_3 I-II=125°; II-III=85.4°; III-IV=35°

tr.t. $AgNO_3$ I-II=159.6°

Flawitzki, 1909

E: 30.94 mol% 102.4°

Ammonium nitrate ($N_2H_4O_3$) + Calcium nitrate (CaN_2O_6)

Urbanski and Kolodziejczyk, 1936

%	f.t.	m.t.	E
0	169.5	-	-
3	-	140	-
5	162.5	128	-
10	153.5	115	-
15	141.0	-	106.2
20	127.0	-	110.3
25	117.1	-	111.2
27.5	113.2	-	111.1
30	117.0	-	110.6
32.5	129.5	-	111.0
35	138.8	-	111.0
40	148	-	110.5
80	-	112	-
90	-	155	-

Clark, Clow and al., 1949

mol%	f.t.	mol%	f.t.
0.0	169.5	10.8	135.0
5.6	154.9	12.1	143.2
8.0	147.6	14.8	154.5
9.8	137.4	15.3	158.0
10.0	131.0		

E. 18.8 wt% 129° (sic.)

Ammonium nitrate ($N_2H_4O_3$) + Barium nitrate (BaN_2O_6)

A.N.Campbell and A.J.R.Campbell, 1947

%	f.t.
5	163
7	167
10	189-185

Ammonium nitrate ($N_2H_4O_3$) + Lead nitrate (PbN_2O_6)

Bogitch, 1915

%	f.t.	E	%	f.t.	E
0	158	-	36	142	130.5
9	156	-	40	155	129
17	150.5	-	43	164	124
23	144	-	45.5	180	128
29	136.5	131.0	50	202	-
33	131.5	-	55	226	-

Ammonium nitrate ($N_2H_4O_3$) + Cadmium nitrate (CdN_2O_6)

A.N.Campbell and A.J.R.Campbell, 1947

%	f.t.	%	f.t.
10	159-155	40	96.5-94.5
20	141-137	50	98.5
30	119	60	111.5

Ammonium orthophosphate (NH_6O_4P) + Potassium orthophosphate (KH_6O_4P)

Krickmeyer, 1896

%	t	d
100	20	2.338
5.5	20	1.825
0	19	1.803

Ammonium sulfate($N_2H_8SO_4$) + Sodium sulfate hydrate ($Na_2SO_4.10$ aq.)

Dawson, 1918

%	tr.t.
100	32.35
95.55	30.05
91.46	27.75
84.28	26.50
78.19	26.50

Ammonium sulfate ($N_2H_8O_4S$) + Potassium sulfate (K_2SO_4)

Sommerfeldt, 1899-1901

mol%	Q mix	mol%	Q mix
	mixed crystals		
1.62	5	9.81	56
3.66	7	11.18	123
5.10	7	12.20	132
6.80	11		

by mole mixed crystals

Retgers, 1889

%	d	%	d
94.55	2.574	73.53	2.342
91.67	.578	70.70	.323
84.97	.474	57.33	.187
81.55	.451	34.65	.004
79.45	.432	16.17	1.883

Wulff, 1907

%	angle of optical axes (red light)
	(100)
100	56.12
93.7	41.27
89.6	34.30
87.6	27.20
84.4	17.51
81.4	10.56
81.3	8.33
	(001)
80.6	22.56
80.42	31.06
77.66	49.36
74.71	67.10
71.43	80.46

%	d	Birefringence.10^5
100	-	1228
98.5	2.647	287
97.4	.630	268
95.5	.608	248
76.5	.380	220
22.5	1.919	116
8.5	.825	875
0	-	1100

Ammonium acid sulfate (NH_5O_4S) + Potassium acid sulfate (KHO_4S)

Rogers and Ubbelohde, 1950

%	f.t.
100	207.1
17	110.5 E
0	144.8

Ammonium sulfate ($N_2H_8O_4S$) + Rubidium sulfate (Rb_2SO_4)

Wulff,1907

%	d	Birefringence.10^5 a	b	c
0	1.769	1228	989	240
51.3	2.40	923	780	151
67.0	2.69	746	665	81
68.5	2.72	686	620	92
79.1	2.97	-	513	-
79.5	2.98	549	502	45
100	3.611	95.3	110.5	15.2

Ammonium sulfamate ($N_2H_6O_3S$) + Sodium sulfamate (NaH_2NO_3S)

Laning and Van der Meulen, 1948

%	f.t.	E	%	f.t.
0.0	132.85	-	35.0	168.5
5.0	128.80	-	40.0	177.0
10.0	124.75	-	50.0	193.0
15.0	120.70	118.7	60.0	204.4
16.2	119.80	118.8	65.0	209.9
16.8	119.20	"	70.0	212.3
16.95	118.8	-	72.5	212.9
17.0	119.0	118.8	73.0	212.6
17.5	121.0	118.8	75.0	215.0
18.0	122.9	118.8	77.5	218.0
20.0	130.2	-	80.0	221.5
25.0	144.7	-	100.0	250.5
30.0	157.9	-(2+5)		

Iodine monochloride (ICl) + Potassium chloride (KCl)

Fialkov and Chor, 1949

mol%	f.t.	mol%	f.t.
0.0	27.0	8.73	13.0
0.91	25.5	8.79	13.0
1.01	24.5	10.3	12.5
2.88	22.0	11.84	12.0
3.04	21.5	15.6	13.0
5.15	16.0		

mol%	ϰ		mol%	ϰ	
	35°	45°		35°	45°
0.00	57.0	63.0	8.1	170	210
0.97	93.2	110	9.9	158	190
2.2	124	151	13.2	150	185
4.3	152	180	16.7	150	183
6.2	170	210			

Iodine monochloride (ICl) + Aluminum chloride ($AlCl_3$)

Fialkov and Chor, 1949

mol%	f.t.	E	mol%	f.t.	E
0.0	27.0	-	36.51	110	-
2.49	24.0	-	38.16	110	-
4.87	21.5	-	40.55	108	-
7.82	13.5	-	42.14	105	-
10.59	7.0	5.0	46.56	132	105
13.31	2.0	-	51.61	138	-
13.98	-	4.5	60.49	155	-
15.88	9.0	5.0	69.50	170	-
18.19	23.5	6.0	80.45	178	-
20.30	32.0	5.0	90.71	186	-
28.55	100.0	-		(2+1)	

mol%	d		
	35°	45°	65°
0	3.184	3.159	-
1.14	.167	.131	-
4.49	.141	.106	-
4.44	.067	.042	-
8.11	.025	2.992	-
9.85	2.989	.977	2.932
12.77	.956	.938	.912
17.0	.916	.887	.859
20.38	.871	.839	.798
23.99	.837	.806	.764
27.22	.780	.759	.739
29.13	.748	.728	.708

mol%	ϰ		mol%	ϰ	
	35°	45°		35°	45°
0.0	55.1	56.1	13.22	75.3	96.2
0.81	51.3	53.7	19.65	59.8	80.3
1.42	47.3	52.1	24.77	47.1	64.8
1.79	51.8	57.5	30.80	38.9	53.7
2.41	55.7	65.2	36.66	35.6	50.3
5.89	67.0	86.1	45.04	32.0	44.8
7.28	70.9	91.5	49.08	31.3	42.8
9.44	77.8	97.9	53.71	31.1	41.0

Iodine monochloride (ICl) + Aluminum bromide ($AlBr_3$)

Fialkov, 1940 and 1941

%	ϰ			τ.10²
	35°	45°	100°	
0	52.6	54.9	-	1.02
3.25	38.5	42.4	-	1.96
6.26	48.6	58.1	-	2.52
10.12	58.9	73.8	-	-
13.71	66.4	-	-	2.86
18.68	68.7	88.3	-	-
24.97	60.5	-	-	3.40
29.10	57.8	77.5	-	-
46.69	27.6	-	-	3.64
52.13	21.2	28.9	-	3.97
54.68	11.4	15.9	138.2	5.40
57.71	17.0	26.1	-	5.40
58.31	10.0	-	121.6	-
62.28	-	23.9	-	-
65.14	12.7	19.5	-	-
67.44	-	18.3	-	-
69.46	6.5	12.8	81.9	9.56
78.28	3.7	6.1	35.5	6.35

Iodine monobromide (IBr) + Aluminum bromide ($AlBr_3$)

Fialkov and Chor, 1953

mol%	f.t.	mol%	f.t.	
0	40.0	54.6	78.0	
8.0	31.5	60.0	83.0	
12.3	18.0	66.1	89.0	
44.6	82.0	75.6	89.0	
51.7	85.0	86.9	90.0	
52.7	83.0	100.0	90.0	(1+1)
53.6	82.0			

Fialkov, 1940 and 1941

mol%	d	mol%	d
	100°		
8.78	4.122	69.50	3.493
20.55	4.025	77.70	3.254
29.23	3.948	92.40	2.844
41.42	3.842	94.85	2.785
54.84	3.702	100	2.635
62.90	3.613		

mol%	η	mol%	η
	60°		
8.78	2725	62.90	7941
20.55	4320	69.50	6990
29.53	5430	77.70	4615
41.42	7902	94.85	3184
54.84	9210		

Fialkov , 1940 and 1941

wt%	mol%	κ 45°	κ 55°
0	0	6.4	10.4
0.85	0.65	2.1	-
2.25	1.75	18.0	20.3
4.45	4.40	34.7	-
9.17	7.27	51.5	61.2
10.78	8.56	52.9	-
14.02	11.23	53.9	67.7
18.99	15.40	53.6	67.4
22.30	18.40	48.1	61.7
26.04	21.45	44.1	-
29.56	24.53	43.5	57.2
33.41	28.01	36.0	-
35.67	30.06	33.8	45.9
37.61	31.85	31.8	-
40.66	34.70	29.4	41.4
46.79	36.95	27.7	-
47.34	41.08	25.8	-
54.80	48.45	21.0	-
57.09	50.78	16.3	-
62.96	58.86	9.4	-
71.04	65.54	8.4	-
82.55	78.60	2.0	-

wt%	mol%	κ	wt%	mol%	κ
		100°			
6.09	4.80	90.7	62.69	56.54	66.1
13.78	10.90	144.5	62.96	56.86	67.9
21.52	16.70	142.9	71.04	65.54	42.7
30.07	25.02	123.4	82.55	78.60	6.8
37.76	32.00	112.5	85.35	81.87	2.6
53.12	46.80	91.8			

Sulfur monobromide (S_2Br_2) + Stannic tetrabromide ($SnBr_4$)

Pushin and Makucz, 1938

mol%	f.t.	m.t.	mol%	f.t.	m.t.
0	-46	-	48.5	-4.5	-55
4.5	-	-47.5 E	58.5	+1	-52
14	-37	-	67	8	-
15	-	-48	74.5	14	-
23	-27.5	-48	80	17.5	-60
29.5	-20.5	-48	89.5	23.5	-
34.5	-15.5	-54	100	30	-
42	-10.5	-54			

mol%	f.t.	E	mol%	f.t.	E
100	30	-	40	- 11	-51
90	24	-	30	- 20	-48
80	18	-	20	- 29	-47
70	11	-	10	- 71	-47
60	4	- 51	4.5	- 47.5	-47
50	-3	- 55	0	- 46	-

Bromine trifluoride (BrF_3) + Uranium hexafluoride (UF_6)

Fischer and Vogel, 1954

mol%	f.t.	m.t.	E
0	8.7	-	-
2.6	-	7.1	6.7
5.1	-	-	6.4
6.0	-	16.3	-
7.6	-	-	6.4
8.4	-	-	6.5
8.5	-	27.1	-
9.9	-	-	6.5
13.4	-	38.9	-
14.7	-	-	6.3
19.3	46.5	45.4	-
19.5	-	46.5	-
24.8	48.9	48.8	-
29.1	-	52.2	-
30.2	51.8	51.9	-
35.6	53.3	-	6.1
38.2	-	55.0	-
41.1	55.1	-	6.3
45.4	55.9	-	6.1
48.7	56.5	-	-
51.9	56.7	-	-
53.9	56.9	-	6.3
56.3	57.3	-	-
58.8	57.7	-	6.4
63.1	57.9	58.0	6.2
67.0	58.4	58.4	-
69.5	58.6	-	-
72.9	59.1	59.3	6.0
76.8	59.6	59.6	-
80.2	60.0	-	5.6
81.0	60.0	-	-
81.8	60.0	60.2	-
83.1	60.4	60.6	-
85.7	60.0	-	-
88.8	61.0	60.7	-
90.4	61.4	-	-
93.4	61.7	61.6	-
96.4	62.3	62.8	-
100	64.0	-	-

Stein and Vogel, 1954

mol%	n_D	mol%	n_D
	70°		
0	1.4302	53.03	1.3836
4.87	.4248	64.20	.3758
4.98	.4247	81.81	.3656
15.66	.4131	100	.3580
34.84	.3968		

Phosphorus trichloride (PCl_3) + Stannic chloride ($SnCl_4$)

Pushin and Makucz, 1938

mol%	f.t.	E	mol%	f.t.	E
100	-33	-	35.5	-62	-96
88.5	-36.5	-	30	-65.5	-
78	-41	-	24.5	-70	-94
65.5	-46.5	-	21	-73	-
58.5	-49.5	-97	18	-75	-
48	-54.5	-	15	-78	-93.5
			9	-84.5	-93
			0	-92	-

Phosphorus tribromide (PBr_3) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100	97.1	68.5	67.3
92.7	91.3	56.0	53.6
85.6	84.1	51.8	47.0
77.6	76.6	50.9	45.4

Pushin and Makucz, 1938

mol%	f t.	mol%	f.t.	E
100	97	33.3	-	-41.5
92	92	19	-	-40
80	80.5	12.5	-	-40
71.5	71	7	-	-40
64.5	65	4.5	-	-40
55	52	0	-40	-

Phosphorus tribromide (PBr_3) + Stannic bromide ($SnBr_4$)

Pushin and Makucz, 1938

mol%	f.t.	E	mol%	f.t.	E
100	30	-	42	-17	-
90	22.5	-	32	-29	-51
80	15	-	24	-38	-54
69	7	-	16	-51	-50.0 E
60	0	-	9	-45	-54
54	-5.5	-50	0	-40	-
49	-10	-			

Arsenic trichloride ($AsCl_3$) + Germanium tetrachloride ($GeCl_4$)

Green and Kafalas, 1955

mol % L	mol % V	t
97.3	85.6	125.8
91.4	63.5	117.4
85.8	51.4	111.5
66.9	28.2	99.9
57.5	22.7	96.8
42.8	16.1	93.7
33.9	13.5	91.8
28.0	10.7	90.2
15.1	5.61	87.2
9.84	3.93	85.6
8.25	3.34	85.3
6.05	2.51	84.8
5.81	2.34	84.8
1.33	0.524	84.1
1.16	0.476	84.1
0.605	0.255	83.6
0.136	0.051	83.6
0.0136	0.0054	83.6
0.00096	0.00040	83.6
0.00084	0.00032	83.6

Sisler, Pfahler and Wilson, 1948 (fig.)

mol%	f.t. I	mol%	f.t. I	f.t. II
0	-18	80	-46	-
20	-26	88.5	-55 E	-
40	-31	89	-54.5	-55.5 E
60	-36.5	100	-49	-52

Arsenic trichloride ($AsCl_3$) + Stannic chloride ($SnCl_4$)

Sisler, Pfahler and Wilson, 1948 (fig.)

mol%	f.t.	mol%	f.t.
0	-18	69	-48 E
20	-25	80	-43.5
40	-34	100	-34
60	-43.5		

Arsenic tribromide ($AsBr_3$) + Zinc bromide.Aluminum bromide ($ZnAlBr_5$)

Gorenbein and Kriss, 1952 fig.

%	d	η	κ
	110°		
0	3.168	1243	-
32.1	.089	2385	0.51
45.1	.062	3258	1.44
55.7	.042	4929	3.06
76.0	.012	15328	7.05
84.3	.000	27484	7.92
93.0	2.992	57500	7.82
94.4	.998	70405	7.78
96.7	.987	89068	7.55
100.0	.986	124363	7.21

%	κ 100°	120°	130°	140° (fig.)
32	0.4	0.6	0.7	0.8
40	1.0	1.3	1.4	1.6
50	1.8	2.5	2.7	3.2
60	3.5	4.7	5.6	6.3
70	5.2	7.7	9.1	10.7
80	6.2	9.8	11.8	13.9
90	6.0	10.2	12.9	15.4
95	5.3	10.0	12.9	16.0
100	4.9	9.8	12.2	15.4

Arsenic tribromide ($AsBr_3$) + Aluminum bromide ($AlBr_3$)

Kendall, Crittenden and Miller, 1923

mol%	f.t.	mol%	f.t.
100.0	97.1	53.6	60.9
94.1	94.5	43.1	52.8
80.7	81.1	32.0	41.7
67.2	69.5	19.1	28.2
59.9	65.7	3.9	32.8
55.2	61.9	0.0	32.8

Izbekov, 1925

mol%	f.t.	mol%	f.t.
100	97.4	51.1	55.0
94.3	92.0	42.0	47.6
91.0	88.4	33.9	40.9
85.3	82.5	27.5	36.1
80.5	78.1	15.0	28.0
73.4	73.0	10.9	29.1
69.5	69.4	7.8	30.2
64.3	65.3	0.0	32.0
59.2	61.2		

Pushin and Makucz, 1938

mol%	f.t.	E	mol%	f.t.	E
100	97	-	28	25.5	25.5
90	87.5	-	20	26.3	24
80	79	-	18.5	-	25.5
70	71	23	17.5	27	-
60	61	25	10.5	28.5	21
50	51.5	25.5	3.5	30.2	-
40	41	25.5	0	31	-
33.5	31	25			

Arsenic tribromide ($AsBr_3$) + Antimony tribromide.Aluminum bromide ($SbBr_3.AlBr_3$)

Gorenbein and Kriss, 1951

%	d	η	κ
	85°		
0	3.232	1518	-
16.26	.231	1785	8.9
32.86	.253	2381	12.1
41.41	.264	2873	25.2
51.94	.276	3745	45.6
59.92	.286	4668	60.9
67.85	.304	6115	76.6
77.31	.310	8422	85.5
86.83	.330	12623	89.2
95.40	.341	18987	85.7
100.00	.346	23466	80.3
	90°		
0	3.219	1467	-
16.26	.217	1715	8.7
32.86	.239	2228	12.0
41.41	.251	2659	25.7
51.94	.265	3399	47.8
59.92	.274	4191	64.9
67.85	.291	5468	82.4
77.31	.298	7440	94.0
86.83	.319	10805	101.5
95.40	.330	15984	98.7
100.00	.335	19561	91.5
	95°		
0	3.206	1407	-
16.26	.204	1664	8.5
32.86	.225	2098	11.7
41.41	.238	2496	26.1
51.94	.251	3139	49.8
59.92	.261	3840	69.4
67.85	.277	4975	88.4
77.31	.286	6608	103.6
86.83	.307	9439	112.3
95.40	.318	13715	111.0
100.00	.322	16399	102.5
	100°		
0	3.193	1328	-
16.26	.191	1624	8.2
32.86	.210	1979	11.6
41.41	.225	2360	26.4
51.94	.237	2912	51.4
59.92	.247	3546	72.5
67.85	.263	4485	93.8
77.31	.274	5880	112.5
86.83	.295	8282	122.8
95.40	.307	11780	123.2
100.00	.313	13957	115.5

Arsenic tribromide ($AsBr_3$) + Bismuth tribromide ($BiBr_3$)

Pushin and Makucz, 1938

mol%	f.t.	mol%	f.t.	E
100	218	50	191.2	30.3
89.5	210	34.5	193.5	-
79.5	202	23	188.5	-
68	197	13	175	-
52	195	4	136	-
		0	31	-

Arsenic tribromide ($AsBr_3$) + Stannic bromide ($SnBr_4$)

Pushin and Makucz, 1938

mol%	f.t.	E	mol%	f.t.	E
0	31	-	55.5	3.5	3.5
7.3	27	-	62.6	6.5	0.2
15.2	23.2	1.0	68.4	9.2	0
23.5	19	2.0	73.2	11.1	-0.2
32.4	14.5	2.0	80.4	16.3	-1.5
41.8	10.0	2.5	86.8	20.5	-3.0
51.9	4.0	2.5	100	30	-
55.0	3.5	-			

Arsenic triiodide (AsI_3) + Aluminum iodide (AlI_3)

Nizhnik , 1937

wt%	mol%	f.t.	E
100	100	188.9	-
93.19	94.10	174.0	116.8
86.64	87.76	162.4	116.8
79.81	82.18	158.0	116.4
69.46	72.72	153.8	119.0
60.75	64.35	146.0	116.8
51.59	55.46	135.7	117.5
46.53	54.08	129.0	-
46.96	50.81	128.8	116.4
44.06	47.86	127.1	116.3
41.74	44.22	122.6	-
39.75	43.48	120.7	117.0
37.97	41.61	120.0	118.0
35.65	39.25	121.3	116.7
22.84	25.67	127.2	118.0
14.36	16.36	131.4	116.9
0.0	0.0	138.8	-

Boron tribromide (BBr_3) + Aluminium tribromide ($AlBr_3$)

Adamsky and Wheeler, Jr., 1954 (fig.)

mol%	f.t.	mol%	f.t.
0	-42	50	45
10.48	-46.1	70	68
20	0	80	80
30	18	100	98
40	36		

Boron tribromide (BBr_3) + Stannic bromide ($SnBr_4$)

Adamsky and Wheeler, Jr., 1954 (fig.)

mol%	f.t.	mol%	f.t.
0	-48	70	0
18.3	-52.5	90	+18
30	-39	100	29
50	-18		

Boron tribromide (BBr_3) + Stannic iodide (SnI_4)

Adamsky and Wheeler, Jr., 1954 (fig.)

mol%	f.t.	mol%	f.t.
0	-40	60	105
0.94	-51.1	70	115
10	+42	80	121
20	68	90	132
30	82	100	144
50	100		

Bromine pentafluoride (BrF_5) + Uranium hexafluoride (UF_6)

Stein and Vogel, 1954

mol%	n_D	mol%	n_D
	70°		
0	1.3275	59.65	1.3477
16.84	1.3338	76.07	1.3522
31.53	1.3388	90.69	1.3560
43.54	1.3428	100	1.3580
51.34	1.3449		

Phosphorous pentachloride (PCl_5) + Aluminum chloride ($AlCl_3$)

Fialkov and Buryanov, 1953 (fig.)

mol%	f.t.	mol%	f.t.
100	195	50	345 (1+1)
80	150	45	320
67	52	0	160

Phosphorous pentachloride (PCl_5) + Ferric chloride ($FeCl_3$)

Fialkov and Buryanov, 1953

mol%	f.t.	mol%	f.t.
100	300	50	332 (1+1)
80	240	45	320
67	98	0	160

Silicium tetrachloride ($SiCl_4$) + Stannic iodide (SnI_4)

Hildebrand and Negishi, 1937

mol %	f.t.	mol %	f.t.
0.155	0.2	32.53	139.4 sat.t.
0.381	25.0	37.62	139.8 "
0.639	40.0	43.16	139.7 "
2.36	81.3	43.54	139.5 "
6.43	112.1	45.06	139.1 "
7.40	115.6	46.95	138.8 "
15.36	131.0	51.21	135.7 "
21.56	130.9 sat.t.	51.62	135.8 "
24.21	133.4 "	55.12	132.4 "
29.59	138.2 "		

T.C.D. 39 mol % 139.9°

Silicium tetrachloride ($SiCl_4$) + Titanium tetrachloride ($TiCl_4$)

Nasu, 1933

mol%	f.t.	mol%	f.t.	E
100.0	24.8	29.4	55.2	-
95.4	27.0	23.8	59.3	-
88.1	30.3	17.2	63.3	-
83.8	31.0	9.4	65.3	-
77.6	33.4	7.7	66.8	70.3
69.4	37.6	5.9	68.3	"
63.4	40.4	4.0	68.3	"
56.5	43.9	3.0	69.3	"
51.0	46.7	2.0	69.3	"
48.7	48.7	0.0	70.4	"
38.5	52.6			

mol%	m.t.	E	mol%	m.t.	E
89.7	28.0	67.3	45.5	50.4	70.3
81.3	31.4	70.3	34.2	54.2	"
75.4	34.6	69.3	23.8	58.9	"
59.8	42.0	70.3	17.2	60.3	70.9
53.7	45.3	70.3	5.9	66.6	"

Selenium tetrachloride ($SeCl_4$) + Aluminum chloride ($AlCl_3$)

Houtgraaf and al; 1953 (fig.)

mol%	f.t.	mol%	f.t.
100.0	196.3	64.12	111.3
90.16	195.0	60.40	130.2
80.10	184.7	56.22	151.0
71.97	155.1	52.97	159.0
67.36	120.8	51.73	160.5
66.45	112.9	50.73	162.3
66.07	111.0	50.0	163.8
65.41	106.5	0.0	305 ± 5

Tellurium tetrachloride ($TeCl_4$) + Aluminum chloride ($AlCl_3$)

Houtgraaf and al; 1953

mol%	f.t.	mol%	f.t.	
100.0	196.0	47.72	141.8	
89.71	192.6	47.0	139.7	
80.0	183.4	44.50	133.9	
69.35	155.9	42.32	128.2	
65.43	132.5	40.34	120.6	
63.13	115.7	37.80	125.0	
61.01	117.7	36.37	135.5	
58.20	129.8	31.39	160.1	(1+1)
53.95	143.5	20.87	196.9	
50.0	149.3	10.46	217.7	
48.93	144.7	0.0	227.9	

Boron oxide (B_2O_3) + Lithium oxide (Li_2O)

Mazzetti and de Carli, 1926

%	f.t.	%	f.t.
32.00	722	17.00	904
29.68	760	15.74	810
29.35	755	14.80	655
27.03	750	13.50	750
26.00	720	12.00	700
22.00	680	10.40	730
20.00	730	9.00	620
18.80	850	7.70	680
18.20	925	7.00	590
		7.00	520
		5.60	490

(1+1) (2+1) (3+1) (4+1) (5+1)

Rollet and Bonaziz, 1955

mol%	f.t.	
69.5	695	tr.t. Li_2O + (1+3)
66	630	E (1+3) + (2+3)
65	610	E metastable
62.5	685	tr.t.(2+3) + (1+1)
50	840	(1+1)
47.5	833	E (1+1) + (2+1)
33.3	915	(2+1)
18	840	tr.t.(2+1) + (3+1)
12	790	tr.t.(3+1) + (4+1)
7	730	tr.t.(4+1) + x

Bresker and Evstropiev, 1952

wt%	mol%	d	n_D
		20°	
0	0	1.8343	1.4625
0.92	2.13	.8621	.4687
1.82	4.14	.8827	.4723
2.75	6.18	.9145	.4810
3.69	8.16	.9457	.4867
4.68	10.23	.9772	.4937
5.66	12.24	2.0094	.5009
6.65	14.21	.0381	.5067
7.75	16.35	.0682	.5136
8.68	18.16	.0941	.5188
9.85	20.27	.1228	.5251
10.95	22.26	.1506	.5307
11.99	24.09	.1768	.5364
13.09	25.98	.1999	.5421

Boron oxide (B_2O_3) + Sodium oxide (Na_2O)

Burgess and Holt, 1905

mol %	f.t.	softening t (of glass)	mol %	f.t.	softening t (of glass)
0	-	468	36.00	-	613
2.44	-	470	38.46	777	620
5.88	-	528	42.86	815	610
20.00	858	628	50.00	930	615
33.33	791	628	55.56	960	-

Ponomareff, 1914

%	f.t.	m.t.	%	f.t.	m.t.
30.8	732	730	18.3	783	779
29.3	637	630	18.0	760	-
28.1	580	574	17.7	764	754
26.2	615	-	17.1	749	732
24.7	679	672	16.2	719	719
22.8	694	688	15.3	670	664
22.3	680	667	14.0	607	600
21.7	658	655	12.5	582	574
20.2	709	-	10.9	585	579
19.3	746	739			

(3+1) (4+1)

Jenckel, 1936

%	f.t. 1	f.t. 2	f.t. 3	f.t. 4
12.10	738 (4+1)	-	-	-
12.45	753 I	733 (4+1)	-	-
14.80	782	771 II	-	-
17.01	820	795	-	-
18.51	820	797	-	-
19.35	816	794	766 (3+1)	716 (3+1)
20.50	810	780	770 I	720 II
21.70	-	-	770	728
22.74	-	-	769	724
23.74	-	-	769	718
24.05	-	-	765	719
25.08	-	-	760	714
26.22	-	-	740	694
26.89	-	-	729	684
27.13	-	-	729	-

%	f.t. 5	f.t. 6	f.t. 7	tr.t.
25.08	725 (2+1)	692	-	-
26.22	723 I	678 II	-	-
27.13	-	682	-	-
27.74	729	681	-	-
29.51	732	664	-	-
30.04	735	-	-	-
30.78	742	664	-	-
31.23	-	-	-	740
32.05	740	-	759-769(1+1)	740
32.61	738	666	-	-
35.07	-	-	836	738
41.28	-	-	916	737
44.70	-	-	948	737
47.10	-	-	974	-

Morey and Merwin, 1936

%	f.t.	
47.21	966	(1+1)
36.64	839	
32.5	744	
30.77	742	(2+1)
30.01	739	
29.82	737	
27.88	723	
27.46	719	(2+1) + (3+1)
	725	(3+1)
26.88	736	
25.07	759	
24.76	761	
23.60	766	(3+1) + (4+1)
	769	(4+1)
23.08	766	(3+1) + (4+1)
	776	(4+1)
22.74	766	(3+1) + (4+1)
	777	(4+1)
21.68	792	
20.09	803	
18.62	814	
18.20	815	
17.78	814	
15.5	800	
11.75	753	
7.01	661	
5.24	620	
2.99	585	
1	560	

Leontyeva, 1937

%	tw	%	tw	
0	250	13.9	428	
2.4	300	15.4	487	(1+3)
5.6	425	16.9	435	
7.7	403	22.1	443	
10.5	410	100	438	

tw = softening t.(inflexion point between f.t. and m.t.)

Wolarovitsch, 1935

%	d 600°	800°	1000°	1200°
33.5	2.160	2.028	1.923	1.835
22.8	2.041	1.961	.876	.795
18.3	-	.894	.832	.770
6.2	1.692	.629	.590	.565
0	1.529	.475	.445	.429

t	d	η	t	d	η
		6.2%			
1078	1.5723	8.33	806	1.6313	65.79
1330	.5504	-	682	.6597	413
1256	.5564	-	560	.7065	8700
1107	.5773	7.435	488	.7421	-
937	.6077	19.80			
		22.8%			
1326	1.7482	-	838	1.9436	31.3
1297	.7606	-	833	.9474	38.2
1208	.7953	-	794	.9666	80.32
1201	.7953	-	781	.9704	110
1193	.8018	-	750	.9841	263
1117	.8298	-	737	.9881	398
1103	.8198	-	730	.9920	490
1081	.8450	-	688	2.0100	2450
1004	.8762	1.7	686	.0060	2670
993	.8709	4.31	672	.0121	5180
966	.8904	-	631	.0305	4410
924	.9157	8.55	589	.0747	6410
898	.8797	11.8	573	.0619	-
872	.9342	17.5	502	.1277	-
863	.9342	19.8	493	.1277	-
860	.9361	20.8			
		18.3%			
1206	1.7605	-	825	1.8815	46.84
1120	.7985	2.352	799	.8939	100.8
1025	.8215	4.008	760	.9011	213.2
932	.8450	9.242	717	.9194	778.2

Boron oxide (B_2O_3) + Potassium oxide (K_2O)

Rollet, 1935

%	f.t.	
58	950	(1+1)
45.3	787	E (1+1) + (2+1)
40.5	815	(2+1)
37.5	770	E (2+1) + (3+1)
36.5	752	E (2+1) + (4+1) metastable
33.5	825	tr.t. (3+1) + (4+1)
25	857	(4+1)
21.3	780	tr.t. (4+1) + (5+1)

Leontyeva, 1936

t	d	t	d
	33.5%		
1250	1.770	950	1.859
1200	.785	900	.880
1150	.799	870	.887
1100	.818	850	.898
1040	.835	800	.923
1000	.842		

t	d 30.4%	26.8%
1220	1.770	1.764-1.767
1200	.776	.785
1100	.802	.789
1000	.828	.812
900	.855	.828-1.832
800	.869	.855-1.866
700	.919	.916-1.919

t	d 25.8%	d 24.8%
1200	1.739-1.757	1.754
1100	.754-1.776	.776
1000	.785-1.792	.792
900	.821	.815
800	.848	.838-1.848
700	.894	.873
680	-	.830

t	d 16.8%	d 6.7%	t	d 16.8%	d 6.7%
1200	1.666	1.585	800	1.754	1.639
1100	.686	.600	700	.785	.666
1000	.701	.605	680	-	.727
900	.727	.623	560	-	.799

t	d	t	d
	0%		
1287	1.422	892	1.462
1191	.431	773	.475
1101	.437	539	.502
1092	.437	501	.587

Boron oxide (B_2O_3) + Thallium oxide (Tl_2O)

Canneri and Morelli, 1922

mol%	f.t.	E	min.	
29.0	446	-	-	
31.2	466	-	-	
33.3	472	-	-	(2+1)
40.5	448	410	90	
43.1	440	"	180	
45.5	426	"	240	
47.6	-	"	300	
48.4	430	"	240	
50.2	434	-	-	
52.5	"	-	-	
54.8	420	358	-	(1+1)
56.3	414	"	90	
57.7	400	"	120	
59.0	388	"	180	
59.8	-	"	270	
60.3	370	-	-	
62.5	350	-	-	

Boron oxide (B_2O_3) + Magnesium oxide (MgO)

Toropov and Konovalov, 1940

%	f.t.	tr.t. 1	tr.t. 2
10.00	1183	1020	780
15.00	1182	1019	"
20.00	1182	1020	"
22.40	1183	1019	"
25.70	"	1020	"
27.80	"	1019	"
30.00	"	1020	"
35.00	"	1021	"
36.63	1191	1019	-
39.23	1265(1191)	1020	-
44.94	1320(1192)	1020	-
48.85	1354(1192)	1020	-
50.86	1375(1191)	-	-
53.25	1381	-	-
54.49	1365	-	-
55.67	1340	-	-
58.38	1343	-	-
60.70	1383	-	-
63.53	1400	-	-
65.72	1395	-	-
69.30	1390	-	-
70.79	1385	-	-
73.00	1390	-	-

Compl. f.t. : 1191° (1+1)
f.t. : 1381°± 5° (1+2)
f.t. : 1400°± 5° (1+3)

E 1: 56.8% 1300° ± 5°
2: 72.6% 1360° ± 5°

Boron oxide (B_2O_3) + Calcium oxide (CaO)

Carlson, 1932 (fig.)

%	f.t.	E	
4	-	975	(2+1) +L
10	-	"	"
20	-	"	"
23.5	-	"	"
25	-	986	"
28	986	"	(2+1)
30	1015	"	"
40	1140	"	(1+1)
44.5	1155	1125	"
50	1145	"	(1+1) + (1+2)
52	1125	"	(1+1) + (1+2)
60	1290	"	(1+1) + (1+2)
62	1300	1290	(1+2)
64	1290	"	(1+2) + (1+3)
65	1350	"	(1+2) + (1+3)
70	1479	"	(1+2) + (1+3)
71	"	1469	(1+3)
73.5	1469	"	
80	-	"	
90	-	"	

t	% L_1	L_2
900-950	23.0	0.6
	23.1	0.6
	24.9	0.2
975	-	0.7
1100	-	1.5
1200	-	0.9
1300	24.2	1.5
1400	23.5	1.4
1450	23.9	0.9
1500	23.7	1.7

Shermer, 1956

mol %	d at room temp.	t a	b
0	1.859	260	305
29.3	2.484	633	658
	reannealed	646	662
31.4	2.531	647	665
	reannealed	649	665
34.9	2.595	648	668
	reannealed	656	668
	2.633	643	669
	reannealed	657	669

a = critical t. b = deformation t.

mol%	$\tau . 10^6$ 2-100°	100-200°	200-300°	300-400°	400-500°	500-600°
0	16.16	13.77	-	-	-	-
	5.64	6.08	6.54	7.16	7.69	8.68
29.3	5.49	6.02	6.39	7.13	7.71	8.09
	5.73	6.51	6.74	7.67	7.92	8.85
31.4	5.82	6.35	6.81	7.65	8.10	8.31
	6.01	6.76	7.52	7.78	8.38	9.18
34.9	6.62	6.75	7.47	7.85	9.50	9.21
	6.40	7.04	7.46	8.22	8.69	9.55
37.1	6.31	6.84	7.58	8.16	8.76	9.35

Boron oxide (B_2O_3) + Strontium oxide (SrO)

Shermer, 1956

mol%	d at room temp.	t a	b
0	1.859	260	305
21.4	2.649	598	614
21.8	.677	604	621
26.0	.802	619	635
26.9	.852	620	638
28.5	.931	624	641
30.5	3.007	628	649
33.3	.101	629	649

a= critical temperature

b= deformation temperature

mol%	$\tau . 10^6$ 20-100°	100-200°	200-300°	300-400°	400-500°	500-600°
0	16.16	13.77	-	-	-	-
21.4	5.79	6.22	6.66	7.28	8.04	-
21.8	5.77	6.36	6.76	7.40	7.87	-
26.0	5.85	6.46	7.03	7.76	8.68	-
26.9	6.01	6.58	6.86	7.56	8.12	9.04
"	5.97	6.56	7.12	7.67	8.15	9.08
28.5	6.34	6.76	7.24	7.98	8.42	9.42
30.5	6.71	6.96	7.72	8.28	9.26	9.03
33.3	6.96	7.47	8.32	8.49	8.62	10.16

Boron oxide (B_2O_3) + Barium oxide (BaO)

Levin and Mc Murdie, 1949

%	f.t.		n (figure)
below 30.2°	-	$L_1 + L_2$	-
30.2	868	-	1.54
35.51	879	(4+1)	.555
40.6	859	E	.575
52.41	900	(2+1)	.62
55.7	889	E	.625
68.77	1095	(1+1)	.66
77.9	905	E	.685
86.85	1383	(1+3)	.715
87.3	1370	E	-
100	1923	-	-

Shermer, 1956

mol%	d at room temp.	t a	b
0	1.859	260	305
17.3	2.678	509	540
20.1	.813	561	580
23.7	.996	575	595
27.1	3.190	589	606
29.0	.289	587	606
31.7	.429	598	610
36.0	.630	594	609

a= critical temperature

b= deformation temperature

mol%	$\tau . 10^6$ 20-100°	100-200°	200-300°	300-400°	400-500°
0	16.16	13.77	-	-	-
17.3	6.59	6.69	7.25	7.80	9.77
20.1	6.19	6.59	7.01	7.67	8.39
23.7	6.43	6.64	7.02	7.71	8.21
27.1	6.59	6.90	6.01	8.14	9.36
29.0	6.82	7.46	7.89	9.48	10.39
31.7	7.37	7.64	8.13	8.86	9.58
36.0	7.86	8.48	8.91	9.68	10.74

Boron oxide (B_2O_3) + Lead monoxide (PbO)

Mazzetti and De Carli, 1926

%	f.t.		%	f.t.	
100	870		59.70	475	
84.70	250		58.65	460	
81.70	450		58.24	475	
75.00	500	(1+1)	56.38	520	(5+2)
74.00	485		54.30	490	
68.40	420		52.20	530	(3+1)
66.97	360		50.00	550	
65.70	450		49.50	560	
62.48	570	(2+1)	47.70	480	
62.90	550		45.00	560	
60.15	520				

(sic.)

Geller and Bunting, 1937

%	f.t.		%	f.t.
100	886		81.3	545-580
94.5	655	±5	80.3	642
93.7	560	E	80.3	550-560
93.7549	549		79.5	669
93.7	553	α (1+4)	79.5	545-580
93.2	564		78.5	675
92.7	565		78.5	530-590
92.7	552	β (1+4)	77.1	698
92.2	561		77.1	530-595
92.2	554		76.1	705
91.2	548	β (1+4)	76.1	548
90.1	537	β (1+4)	72.5	738
88.0	493	E	68.3	757
88.0	473		61.8	767
87.3	497	α (1+2)	58.5	763
87.3	475		52.0	758
86.6	517	(4+5)	48.6	754
86.6	498		44.9	746
86.6	472		42.8	750
86.2	518		36.7	744
86.0	525		25.9	742
84.9	535		14.1	742
83.7	542		5.1	700
82.5	583	(2+1)	2.1	610
82.5	545	±5	0.0	294
81.3	623			

%	f.t.	sat.t.
36.7	744	777
30.1	-	785
25.9	742	780
14.1	742	777

Kröger and Lieck, 1955

%	sat.t.
10	770
20	784
30	785
36	783
40	770

Shartsis, Spinner and Smock, 1948

%	σ 600°	σ 700°	σ 800°	σ 900°	σ 1000°
0	-	-	75.9	79.5	83.0
4.9	-	-	75.5	78.2	-
10.1	-	-	75.5	78.4	-
16.6	-	-	74.5	78.2	82.1
21.6	-	-	76.1	79.4	-
27.8	-	-	76.3	80.1	-
31.4	-	-	75.5	80.1	-
34.4	-	-	75.4	78.5	-
39.9	-	-	75.4	79.1	-
44.9	-	-	84.8	89.3	-
47.0	-	-	90.6	93.3	-
49.1	-	-	89.4	91.6	-
51.3	-	-	104.0	103.0	-
56.3	-	-	121.8	116.4	-
60.9	-	-	136.8	129.8	-
66.4	-	159.8	150.6	144.2	-
70.7	172.3	165.2	157.7	152.9	-
75.5	-	166.0	162.6	160.3	-
80.0	167.6	165.7	164.3	163.0	-
82.0	163.6	163.0	162.3	162.0	-
84.1	161.6	162.2	162.4	162.8	-
86.0	159.0	160.0	161.6	162.2	-
88.1	156.8	159.1	160.6	162.2	-
89.8	156.5	155.7	158.4	160.4	-
91.1	151.5	152.7	155.4	157.4	-
94.0	-	-	-	151.0	153.1
96.0	-	-	-	144.5	146.2
97.9	-	-	-	132.3	139.3
100.0	-	-	-	132.0	134.8

Melnikova, Evstropiev and Kuznetsov, 1951

21.4 mol%

t	κ
468	5.5×10^{-11}
388	2.0×10^{-12}
350	3.1×10^{-13}
330	1.3×10^{-13}
310	3.0×10^{-14}

21.5 mol%

t	κ
485	1.1×10^{-11}
445	1.8×10^{-11}
406	3.7×10^{-12}
386	1.5×10^{-12}
328	6.9×10^{-14}
305	2.9×10^{-14}
296	1.8×10^{-14}
275	4.6×10^{-13}

24.1 mol%

t	κ
448	2.8×10^{-11}
408	4.9×10^{-12}
370	8.9×10^{-13}
335	2.1×10^{-13}
290	2.3×10^{-14}

28.0 mol%

t	κ
420	2.4×10^{-11}
400	9.6×10^{-12}
367	2.0×10^{-12}
365	2.3×10^{-12}
348	1.1×10^{-12}
308	1.1×10^{-13}
260	4.9×10^{-15}
235	9.1×10^{-16}

30.8 mol%

t	κ
445	4.3×10^{-11}
420	1.8×10^{-11}
328	2.0×10^{-13}
293	3.4×10^{-14}
254	3.6×10^{-15}

33.3 mol%

t	κ
434	5.0×10^{-11}
396	1.1×10^{-11}
364	2.3×10^{-12}
300	8.3×10^{-14}
270	1.8×10^{-14}

36.4 mol%

t	κ
438	1.0×10^{-10}
378	9.1×10^{-12}
370	7.9×10^{-12}
325	8.1×10^{-13}
308	1.8×10^{-13}
280	5.4×10^{-14}
238	3.6×10^{-15}

39.5 mol%

t	κ
370	1.4×10^{-11}
316	9.3×10^{-13}
310	6.1×10^{-13}
290	1.8×10^{-13}
265	5.4×10^{-14}
245	2.1×10^{-14}

42.4 mol%

t	κ
374	5.2×10^{-11}
356	1.8×10^{-11}
316	3.2×10^{-12}
310	2.4×10^{-12}
290	1.3×10^{-13}
280	6.6×10^{-13}
265	1.3×10^{-13}

44.0 mol%

t	κ
386	1.4×10^{-10}
348	2.1×10^{-11}
310	3.2×10^{-12}
270	3.4×10^{-13}
250	1.1×10^{-13}

44.9 mol%

t	κ
368	8.8×10^{-11}
340	2.2×10^{-11}
290	1.7×10^{-12}
252	1.6×10^{-13}

47.3 mol%

t	κ
370	1.1×10^{-10}
325	1.9×10^{-11}
290	3.3×10^{-12}
260	4.6×10^{-13}
230	7.4×10^{-14}

46.8 mol%

t	κ
350	3.6×10^{-11}
300	3.0×10^{-12}
285	1.8×10^{-12}
262	3.6×10^{-13}
248	1.4×10^{-13}
242	1.1×10^{-13}
225	3.0×10^{-14}
215	2.2×10^{-14}

49.2 mol%

t	κ
350	1.6×10^{-10}
310	2.9×10^{-11}
285	6.3×10^{-12}
256	1.2×10^{-12}
230	2.9×10^{-13}
205	5.9×10^{-14}
196	1.9×10^{-14}

49.6 mol%

t	κ
330	1.3×10^{-10}
300	5.4×10^{-11}
285	2.3×10^{-11}
265	9.2×10^{-12}
250	2.6×10^{-12}
230	2.0×10^{-12}
210	3.3×10^{-13}

50.8 mol%

t	κ
325	1.9×10^{-10}
296	5.1×10^{-11}
265	1.7×10^{-11}
250	5.4×10^{-12}
195	2.1×10^{-13}
175	5.0×10^{-14}

50.0 mol%

t	κ
310	1.2×10^{-10}
295	4.6×10^{-11}
276	2.2×10^{-11}

51.0 mol%

t	κ
276	1.8×10^{-11}
260	8.0×10^{-12}
240	2.4×10^{-12}
224	1.1×10^{-12}
170	2.6×10^{-14}

52.0 mol%

t	κ
305	9.7×10^{-11}
250	6.3×10^{-12}
230	2.2×10^{-12}
206	5.6×10^{-13}
180	8.8×10^{-14}

49.6 mol%

t	κ
356	1.1×10^{-10}
315	3.4×10^{-11}
268	2.8×10^{-12}
275	5.2×10^{-12}
245	5.8×10^{-13}
200	3.6×10^{-14}
180	1.0×10^{-14}

53.6 mol%

t	κ
290	1.1×10^{-10}
240	1.4×10^{-12}
216	1.7×10^{-12}

53.0 mol%

t	κ
345	1.7×10^{-10}
300	2.0×10^{-11}
266	2.6×10^{-12}
255	1.2×10^{-12}
235	4.4×10^{-13}
230	2.6×10^{-13}
215	1.5×10^{-13}
210	8.7×10^{-14}

57.2 mol%

t	κ
296	4.2×10^{-11}
266	1.8×10^{-11}
234	7.0×10^{-12}
210	1.9×10^{-12}
185	3.5×10^{-13}

56.0 mol%

t	κ
320	2.7×10^{-10}
300	1.3×10^{-10}
280	4.5×10^{-11}
252	1.1×10^{-11}
234	4.7×10^{-12}
212	1.4×10^{-12}
176	1.5×10^{-13}
150	2.1×10^{-14}

56.0 mol%

t	κ
295	8.3×10^{-11}
285	5.0×10^{-11}
250	1.3×10^{-11}
230	4.5×10^{-12}
215	1.4×10^{-12}
200	7.5×10^{-13}
190	3.4×10^{-13}

56.4 mol%

t	κ
310	2.1×10^{-10}
280	5.8×10^{-11}
230	3.9×10^{-12}
210	1.1×10^{-12}
200	5.6×10^{-13}
150	1.4×10^{-14}

61.0 mol%

t	κ
285	2.6×10^{-10}
265	7.7×10^{-11}
235	2.1×10^{-11}
230	1.6×10^{-11}
210	4.8×10^{-12}
200	3.2×10^{-12}
196	2.7×10^{-12}

62.0 mol%

t	κ
280	3.2×10^{-10}
250	5.6×10^{-11}
240	3.5×10^{-11}
210	7.1×10^{-12}
186	1.7×10^{-12}
170	4.7×10^{-13}
152	1.9×10^{-13}

67.2 mol%

t	κ
230	7.4×10^{-11}
200	1.5×10^{-11}
188	8.4×10^{-12}
176	4.2×10^{-12}
160	1.7×10^{-12}
148	7.0×10^{-13}

69.0 mol%

t	κ
236	2.0×10^{-10}
215	7.2×10^{-11}
196	2.4×10^{-11}
176	7.4×10^{-12}
162	2.8×10^{-12}
156	2.0×10^{-12}

66.0 mol%

t	κ
215	2.8×10^{-11}
186	4.5×10^{-12}
170	1.5×10^{-12}
134	1.7×10^{-13}

69.0 mol%

t	κ
235	1.7×10^{-10}
205	3.6×10^{-11}
184	1.1×10^{-11}
178	6.8×10^{-12}
162	4.0×10^{-12}
156	2.2×10^{-12}

second sample

Boron oxide (B_2O_3) + Cadmium oxide (CdO)

Mazetti and De Carli, 1926

%	f.t.	%	f.t.
77.00	940	62.50	770
75.00	905	59.50	710
72.00	820	57.80	750
70.00	810	57.00	810
67.70	750	54.70	855
66.00	845	51.00	755
64.50	850	52.00	-
63.25	810		

(1+1) (3+2)

Boron oxide (B_2O_3) + Manganese oxide (MnO)

Mazetti and De Carli, 1926

%	f.t.	%	f.t.
60	800	41	850
56	760	38	900
54	750	35	940
53	770	32	910
50	840	29	860
47	820	25	920
45	800	22	870

(1+1) (2+1) (3+1)

Boron oxide (B_2O_3) + Sodium fluoride (NaF)

Bergmann and Nagornyi, 1943

%	f.t.	m.t.	%	f.t.	m.t.
100	990	-	70.6	905	907
92.7	970	-	65	878	880
90	960	962	60	835	838
85	948	950	55	758	860
80	935	937			

Kruh and Stern, 1956

mol%	d 600°	700°	800°
0.0	1.586	1.567	1.548
2.69	.620	.600	.579
7.65	.710	.685	.659
12.4	.782	.752	.720

mol%	η (in poises) 600°	650°	700°	750°
2.7	1800	700	320	160
6.4	1400	500	220	105
12.1	2400	720	240	80
18.9	3600	800	240	120

Boron oxide (B_2O_3) + Potassium fluoride (KF)

Bergmann and Nagornyi, 1943

%	f.t.	%	f.t.
100	856	70.6	736
92.3	830	64.9	686
85.7	806	58.5	584
77.4	772		

Boron oxide (B_2O_3) + Sodium metaphosphate ($NaPO_3$)

Arndt and Geszler, 1908

%	d	%	d
	900°		
100	2.144	5	1.585
50	2.115	1	.552
25	1.820	0.5	.522
10	1.655	0	.520

Arndt, 1907

%	η (in poises)	λ
	900°	
100	1.5	49.5
50	4.5	16.4
5	47.3	1.55
0.5	110.0	0.67

Nitrosyl chloride ($NOCl$) + Aluminum chloride ($AlCl_3$)

Houtgraaf and De Roos, 1953

mol%	f.t.	mol%	f.t.
0.0	-59.5	61.9	123.5
3.3	-62.5	62.3	131
13.2	-76	66.2	151
17.4	-86	67.4	156
19.8	-68	70.0	165
22.5	-26	72.5	176
26.5	-1	75.1	182.5
30.4	+16	77.5	186.5
30.7	17	79.9	189
31.3	17.5	82.2	192
32.1	46	82.8	193.5
34.0	60	83.8	194
35.2	83	84.4	194 (1+5)
38.0	107	88.7	193.5
42.2	148	90.0	195
43.6	157	92.9	194
50.0	180 (1+1)	93.7	194.5
51.3	175	95.0	195
56.3	138	97.6	194.5
58.5	121	99.0	195.5
59.3	116	100.0	195.8
60.7	120		

Nitrogen pentoxide (N_2O_5) + Lead monoxide (PbO)

Baekeland, 1903

t	p.dissoc. (1+1)
223	6.2
230	6.9
250	11.8
274	32.6
296	78.4
357	514.0
448	1180.0

Nitrogen pentoxide (N_2O_5) + Tantalum pentoxide (Ta_2O_5)

Metzger and Lamme, 1915

%	d (boiling $CHCl_3$=1)
0	3.265
10.20	.316
20.06	.484
30.10	.686
40.05	.908
50.02	4.074
59.98	.925
69.94	5.227
79.95	.530
89.80	.843
100	6.201

Phosphorus pentoxide (P_2O_5) + Calcium oxide (CaO)

Nielsen, 1913

%	f.t.	%	f.t.
61	1870	41.3	1200
54.3	1550	28.3	1020
46.2	1300	0	800
40.2	1250		

Trömel, 1932

%	f.t.	%	f.t.
26	950	54	1730 (1+3)
28	975 (1+1)	56.9	1560 E
32	960 E	57	1570
35	1140	60	1690
40	1280	61	1700 (1+4)
44	1300 (1+2)	62.8	1660 E
47	1280	63	1670
50	1530	66	1720

Phosphorus pentoxide (P_2O_5) + Lead monoxide (PbO)

Amadori, 1917

%	f.t.	E	min.
75.84	824	- (1+2)	-
76.06	-	816	160
77.13	868	"	140
78.20	914	814	90
79.27	940	806	30
79.70	946	- (2+5)	-
80.28	-	940	100
81.17	966	"	70
81.80	986	936	50
82.16	1002	928	20
82.48	1014	- (1+3)	-
83.30	998	966	20
83.50	990	"	50
83.95	980	970	60
84.60	-	"	100
84.92	974	"	70
85.74	976	966	40
86.25	980	- (1+4)	-
87.12	972	836	20
88.44	938	840	50
89.50	916	"	80
90.16	888	"	90
90.86	-	"	110
91.95	852	838	40
92.64	860	- (1+8)	-
93.35	854	814	30
94.35	842	820	90
95.50	-	"	120
97.68	860	814	90
98.76	875	810	60

Krol, 1912 (fig.)

mol%	%	C
66.65	75.92	(1+1)
71.43	79.77	(2+5)
74.98	82.54	(1+3)
79.92	86.25	(1+4)
88.89	92.65	(1+8)

Phosphorus pentoxide (P_2O_5) + Ferric oxide (Fe_2O_3)

Brasseur, 1944

They exists only 2+1
3+1 (α + 3)

Phosphorus oxychloride ($POCl_3$) + Zirconium tetrachloride ($ZrCl_4$)

Sheka and Voitovich, 1956

mol%	f.t.	E	mol%	f.t.	E
0.0	1.3	-	40.18	164.2	-
0.48	0.0	-1.5	41.67	172.0	162.5
1.05	-0.50	-1.6	43.35	177.5	162.8
1.38	1.50	-1.8	47.35	201.9	-
2.17	6.00	-	50.91	203.0	-
2.35	7.80	-	54.32	187.9	175.3
5.68	34.5	-	56.26	178.0	-
9.72	55.5	-	57.15	180.0	175.0
10.98	59.5	-	58.46	205.0	175.7
14.36	84.2	-	60.84	245.0	175.7
19.12	114.6	-	62.66	271.0	-
23.54	142.6	-	65.52	314.0	-
25.00	151.0	-	67.62	329.5	-
29.84	175.8	-	69.21	348.0	-
33.09	184.7	-	72.14	360.0	-
34.70	183.3	-	77.85	381.0	-
36.01	181.5	-	100.0	437	-
38.76	171.5	162.0			

(1+1) (2+1)

Arsenic pentoxide (As_2O_5) + Sodium arsenate (Na_3AsO_4)

Amadori, 1911

%	f.t.	E	min	tr.t.	min
33.26	625	550	-	-	-
35.63	615	562	-	-	-
38.00	-	570	-	-	-
40.38	-	"	-	-	-
42.75	582	"	-	-	-
45.13	606	-	-	-	-
47.51	615	-	-	-	-
49.97	605	564	40	-	-
50.93	-	568	80	-	-
53.68	-	586	180	-	-
58.30	626	568	120	-	-
58.77	"	562	100	-	-
61.39	648	546	70	-	-
62.57	660	528	40	-	-
64.48	676	524	30	-	-
67.56	696	516	10	-	-
69.35	722	-	-	697	150
70.54	740	-	-	697	140
73.19	772	-	-	696	110
73.09	785	-	-	694	100
74.34	806	-	-	696	80
75.27	818	-	-	692	"
75.47	824	-	-	694	70
76.81	838	-	-	690	100
78.36	850	-	-	688	120
80.52	838	820	50	-	-
81.39	834	"	70	-	-
82.69	826	"	90	-	-
83.97	846	"	70	-	-
85.93	942	"	"	-	-
87.01	976	816	60	-	-
89.18	1032	814	50	-	-
91.34	1105	810	"	-	-
93.08	1120	804	30	-	-
94.59	1174	796	20	-	-
96.98	1210	795	10	-	-
100	1260	-	-	-	-

(1+1) (3+1) (2+1) (5+1)

Arsenic pentoxide (As_2O_5) + Potassium arsenate (K_3AsO_4)

Amadori, 1911

%	f.t.	E	min	tr.t.	min
36.90	611	523	-	-	-
39.54	595	556	-	-	-
42.17	-	570	-	-	-
44.81	-	"	-	-	-
47.45	596	558	-	435	-
50.08	620	-	-	442	-
52.72	650	-	-	450	-
"	"	-	-	"	60
55.39	618	530	60	"	"
58.04	578	538	100	"	50
59.84	560	542	180	"	40
65.13	610	"	120	446	30
69.04	685	540	80	444	20
70.11	705	538	40	440	"
71.64	730	"	"	-	-
72.82	798	536	20	-	-
73.60	810	530	10	-	-
75.32	872	-	-	-	-
75.89	888	-	-	392	20
76.99	910	-	-	395	30
79.12	948	-	-	398	50
80.20	980	-	-	400	70
81.71	996	-	-	-	-
82.86	990	962	20	400	80
84.44	978	964	50	-	70
86.27	-	"	90	-	50
87.48	1010	"	90	-	40
88.64	1048	"	80	-	20
90.40	1092	"	80	-	20
91.46	1136	"	70	-	10
93.55	1190	"	40	-	-
96.70	1235	"	20	-	-
97.07	1250	"	20	-	-
100	1310	-	-	-	-

(1+1) (3+1) (2+1) and (5+3)

Arsenic pentoxide (As_2O_5) + Lead monoxide (PbO)

Amadori, 1917

%	f.t.		E	min	tr.t.	min
65.96	802		-	-	-	-
66.85	-	(1+2)	792	180	-	-
68.04	842		792	160	-	-
69.40	904		790	120	-	-
71.00	962		786	100	-	-
71.78	998		780	70	-	-
72.80	1030		778	50	-	-
74.40	1042	(1+3)	-	-	-	-
75.32	1035		805	20	834	20
75.86	1030		806	30	"	30
77.34	992		808	50	"	"
78.24	968		810	60	"	40
79.50	937		"	70	"	50
80.28	900		"	100	"	40
81.30	874		813	130	"	10
83.00	-		815	"	-	-
84.70	830		"	"	-	-
86.60	848		"	60	-	-
88.58	862	(1+8)	-	-	-	-
89.80	854		800	40	-	-
91.50	842		804	80	-	-
93.30	825		"	120	-	-
94.90	-		"	150	-	-
96.20	828		802	110	-	-
98.30	860		798	50	-	-
100.00	890		-	-	-	-

Sulfur dioxide (SO_2) + Lithium iodide (LiI)

Ephraim, 1919

t	p	
-20.5	270	(2+1)
-14	380	
-11.5	420	
- 7.5	510	
-12	250	(1+1)
- 9.5	275	
- 7.5	290	
0	450	

Sulfur dioxide (SO_2) + Sodium iodide (NaI)

Ephraim and Kornblum, 1916

t	p	
-20	73	(2+1)
-11.5	110	
-7.5	180	
-5	210	
0	307	
+3	380	
7	478	
11.5	615	
15	750	
-2	510	(4+1)
0	580	

Foote and Fleischer, 1931

I C + L + V

t	p		t	p
-20.20	472	(4+1)	-7.50	840
-15.75	584		-5.10	929
-13.10	658		-3.20	1008
-10.70	728		0.0	1158

II C + C_1 + V

t	p	
-18.65	299	(4+1) + (8+3)
-14.15	424	
-9.20	601	
-7.00	701	
-4.00	852	
0.0	1119	

III S_1 + S_2 + V

t	p	
-24.35	44	(8+3) + NaI
-10.90	133	
0.0	286	
9.90	559	
18.40	964	
19.40	1024	

IV S + S_1 + L + V

t	p	
1.5	1230	(4+1) + (8+3)

Sulfur dioxide (SO_2) + Sodium thiocyanate (NaCNS)

Foote and Fleischer, 1932

C + L + V

t	p	
-21.40	445	(2+1)
-12.20	692	
-2.75	1035	
0.00	1159	

$C + C_1 + V$

t	p	
-21.30	91	(2+1) + NaCNS
-9.70	232	
0.00	499	
+7.00	798	
+10.50	1020	

Sulfur dioxide (SO_2) + Potassium bromide (KBr)

Franklin, 1911

N	λ				
	-33.5	-20	-10	0	+10
2.5246	21.25	28.42	32.78	-	-
1.6420	29.18	34.82	39.08	42.01	-
0.8313	32.19	36.23	39.02	39.88	-
.4209	32.03	33.88	34.92	34.98	-
.2181	31.40	32.22	32.28	31.33	29.92
.1079	31.38	31.08	30.38	29.13	27.32
.0546	33.45	32.65	31.50	29.95	27.55
.02830	37.34	36.32	34.20	32.13	29.30
.02418	38.47	37.55	35.51	33.64	30.73
.01956	-	-	37.57	-	-
.01224	46.11	44.84	42.40	40.25	37.08
.009900	-	-	45.54	-	-
.006200	55.89	55.00	52.85	49.87	45.83
.003138	68.58	67.27	65.88	63.06	57.95
.001589	84.06	84.08	82.74	79.04	73.23
.0008045	100.6	103.9	102.3	99.45	94.40
.0004072	118.7	122.8	125.3	123.8	118.0
.0002062	137.2	144.0	149.3	150.5	147.8

Sulfur dioxide (SO_2) + Potassium iodide (KI)

Walden and Centnertzwer, 1903

mol%	p	mol%	p
0°		-8.1°	
14.74	937	15.18	637
15.97	868	15.63	594
17.16	800	16.03	567
18.71	694	16.44	542
19.72	623	17.50	617
21.45	514	19.16	604
23.72	509	20.22	574
25.23	501		
25.84	504		
26.24	506		
27.83	508		
29.53	505		
33.16	498		
33.55	494		
37.32	491		
41.17	498		
41.33	483		
47.07	501		
48.77	500		
51.89	494		
68.05	482		

Ephraim, 1919

t	p	t	p	
-21	111	-14.5	275	sat.sol.
-12	216	-7	400	
0	483	-2.5	510	
1.5	517			
5	605			
5.5	635			
6.5	680			

Foote and Fleischer, 1931

C + L + V			$C + C_1 + V$		
t	p		t	p	
-18.7	449	(4+1)	-18.20	143	(4+1) + KI
17.05	478		16.05	163	
15.25	512		13.30	199	
13.90	535		9.85	257	
12.25	563		6.50	323	
10.80	587		1.50	457	
9.60	608		0.00	502	
8.50	625				
7.55	637		$C_1 + L + V$		
5.85	652				
4.80	658				
2.95	657		0.95	531	
2.00	648		2.85	589	
-1.00	627		5.85	694	
0.00	606		7.80	770	
			10.00	864	
			13.30	1024	

$C + C_1 + L + V$ 0.2°

Walden, 1899

g/100cc	D b.t.
3.56	+0.12
5.03	+0.15
12.21	+0.380
34.07	+1.215
40.80	+1.895
49.90	+3.045

Walden and Centnerszwer, 1902

N	D b.t.	N	D b.t.
0.371	+0.225	2.05	+1.215
0.735	.380	2.46	1.895
1.13	.550	3.01	3.045
1.58	.835		

Walden and Centnerszwer, 1903

mol%	f.t.	mol%	f.t.
14.51	-11.73	0	-72.7
14.69	-10.30	0.336	-73.0
14.78	-10.86	1.406	-57.4
15.35	-9.23	2.576	-48.6
15.56	-8.63	2.671	-43.2
15.96	-7.59	3.258	-43.3
16.82	-5.61	4.130	-37.3
16.94	-5.30	4.340	-38.4
17.33	-4.59	5.692	-33.1
17.64	-3.80	7.003	-23.4
18.11	-3.10	8.476	-28.4
18.83	-2.26	11.60	-19.2
18.94	-1.77	14.37	-10.9
19.68	-0.70	17.93	-2.5
19.83	-0.50	18.93	-1.2
19.95	-0.35	20.63	+1.0
19.98	-0.24	23.75	+1.8
20.02	-0.42	27.64	+1.8
20.36	-0.11	31.34	+1.8
20.47	+0.14		
20.65	+0.14		
21.12	+0.31		
21.78	+0.37	second serie	
21.82	+0.33		
22.10	+0.30		
23.34	+0.30		
25.32	+0.18		
27.59	+0.24		

Walden and Centnerswer, 1903

%	sat.t.	%	f.t.
1.494	-	1.494	96.4
3.465	86.9	3.465	87.9
5.207	82.3	5.207	86.9
5.630	81.6	5.630	28.7
6.981	79.8	6.981	87.0
9.309	77.5	9.309	89.5
15.73	78.8	15.73	89.6
16.49	78.4	16.49	87.4
24.47	89.3	24.77	84.2
30.87	-	30.87	78.9
		32.48	75.3
		42.23	-

% L_1	% L_2	sat.t.
12	12	77.3 C.S.T.
8.5	15.0	78
6.8	17.6	80
5.1	19.8	82
4.5	21.1	84
3.8	22.5	86
2.7	24.0	88

Walden and Centnerszwer, 1902

N	λ	N	λ
	0°		
0.1578	44.46	0.1664	47.57
.2082	45.22	.2404	44.03
.2887	46.07	.3206	44.83
.4228	46.33	.4373	48.42
.7553	48.22	.6954	50.77
2.2624	37.96	2.0619	43.14

Franklin, 1911

N	λ -33.5°	-20°	-10°	0°	+10°
3.3557	-	-	21.93	26.85	-
2.8225	-	23.88	29.39	34.82	40.44
.2371	24.80	-	35.87	-	-
.1267	-	32.57	37.81	42.60	-
1.8328	29.01	36.72	41.71	46.78	51.82
.7065	29.77	-	-	-	-
.6548	-	37.64	-	-	-
.4144	33.39	-	-	-	-
.3866	-	-	45.01	-	-
.2671	34.70	-	-	-	-
0.9921	37.47	-	-	-	-
.9189	-	44.16	-	-	-
.8913	38.37	44.60	47.07	50.99	54.00
.6246	39.74	-	-	-	-
.5797	39.89	43.82	47.19	49.11	50.11
.2780	40.55	43.77	45.19	-	-
.1850	-	-	43.50	-	-
.1440	40.42	42.24	42.66	42.50	41.77
.1407	41.45	43.30	44.35	43.99	43.13
.09363	41.37	42.64	42.73	42.14	40.88
.04742	44.30	45.35	44.84	43.69	42.06
.02400	-	-	50.27	-	-
.02088	-	-	52.05	-	-
.01797	53.99	54.74	53.98	52.10	49.71
.01674	54.40	55.39	54.82	53.32	51.02
.01388	-	-	57.25	-	-
.01358	-	-	57.29	-	-
.01089	60.34	61.46	61.04	59.01	56.62
.009025	-	-	64.83	-	-
.006878	-	-	68.95	-	-
.004589	-	-	77.63	-	-
.004568	-	-	78.18	-	-
.003481	-	-	83.78	-	-
.002320	-	-	94.70	-	-
.002305	-	-	95.94	-	-
.002086	-	-	98.20	-	-
.001762	-	-	101.7	-	-
.001684	96.40	101.5	104.0	102.6	99.67
.001654	96.70	102.8	103.9	-	-
.001456	-	-	108.0	-	-
.001165	-	-	114.3	-	-
.0011646	-	-	113.4	-	-
.001097	106.2	112.8	115.0	115.3	112.4
.001056	-	-	117.9	-	-
.0008929	-	-	122.2	-	-
.0008584	112.2	119.3	124.2	124.5	123.2
.0007369	-	-	129.0	-	-
.0005896	-	-	129.9	-	-
.0005889	-	-	135.7	-	-
.0005348	-	-	138.9	-	-
.0004579	126.0	134.7	144.1	146.3	144.7
.0004517	-	-	144.4	-	-
.0004348	128.2	137.0	145.2	147.3	148.1
.0003731	-	-	147.4	-	-
.0002981	-	-	158.2	-	-
.0002787	-	-	161.2	-	-
.0002287	140.9	155.0	165.1	171.5	173.8
.0002201	142.1	153.7	164.9	170.6	173.3
.0001890	-	-	171.2	-	-
.0001511	-	-	175.1	-	-
.0001158	-	-	190.7	-	-
.0001115	155.8	171.0	185.5	194.3	201.5
.00007651	-	-	192.2	-	-
.00005956	162.3	186.4	200.4	212.8	227.5
.00005643	-	-	204.7	-	-

Sulfur dioxide (SO_2) + Potassium thiocyanate (KCNS)

Ephraim, 1919

t	p		t	p	
-21	85	(1+1)	0	25	(1+2)
0	357		9.5	54	
+5	480		10.5	60	
6.5	515		13.5	78	
12	755		23	148	
13	780		39.5	400	
			41.8	497	
			47	693	
			49	745	

Foote and Fleischer, 1932

C + L + V t	p		C + C_1 + V t	p	
-21.30	445	(2+1)	-10.00	183	(2+1) + (1+2)
14.50	619		0.00	371	
7.50	842		+11.00	738	
0.00	1143		14.50	912	
			16.60	1029	

C_1 + C_2 + V

t	p	
0.00	285	(1+2) + KCNS
12.00	70	
19.00	112	
20.85	129	
24.00	154	

Walden and Centnerszwer, 1899 and 1909

N	D b.t.
1.47	-0.66
2.86	1.68
4.54	5.04

Sulfur dioxide (SO_2) + Rubidium iodide (RbI)

Ephraim, 1919

t	p	t	p	
0	289	10.8	577	
+5	403	12.5	630	(4+1)
9	520	13.5	670	

Foote and Fleischer, 1931

C + L + V			C + C_1 + V			C_1 + L + V	
t	p		t	p		t	p
-18.1	468	(3+1)	-21.3	55	(3+1)	+21.75	1130
-10.3	640		-10.0	134	+ RbI	22.00	1141
-5.6	760		0.0	273.5		23.0	1194
-2.0	857		+6.3	426			
0.0	915		10.90	572			
+5.0	1059		13.30	758			
			19.20	974			

C + C_1 + L + V t= 21.3°

Walden and Centnerszwer, 1902

molarity	D b.t.
0.239	0.156
.395	.259
.562	.332
.759	.422
1.012	.537

Sulfur dioxide (SO_2) + Rubidium thiocyanate (RbCNS)

Ephraim, 1919

t	p	
(1+2)		
0	74	
11.5	164	
19	290	
-15	270	sat. sol.
- 5	400	
0	525	
+3.5	620	

Sulfur dioxide (SO_2) + Cesium iodide (CsI)

Ephraim, 1919

t	p		t	p	
-20.5	70	(4+1)	-11.5	575	sat.sol.
-11.5	130		-10.5	618	
-0.5	250		-9.5	650	
+6	390				
12.5	590				
16	710				

Foote and Fleischer, 1931

t	p		t	p	
-21.60	427	(3+1)	-19.60	63	(3+1) + CsI
-18.45	488		-15.90	85	
-16.30	539		-10.90	123	
-11.50	658		-5.25	184	
-8.10	755		0.0	266.5	
-4.25	871		+7.50	438	
0.0	1025		12.40	603	
			17.10	808	
			20.25	982	

Sulfur dioxide (SO_2) + Cesium thiocyanate (CsCNS)

Ephraim, 1919

t	p	
-15.5	125	(2+1)
0	310	
+9.5	470	
11.5	515	
19	735	

Sulfur dioxide (SO_2) + Calcium iodide (CaI_2)

Ephraim, 1919

t	p		t	p	
0	46	(4+1)	0	47	(4+1)
7	75		7	79	
16	200		17.5	240	
33	760		27.5	485	

Sulfur dioxide (SO_2) + Calcium thiocyanate ($CaC_2N_2S_2$)

Ephraim, 1919

t	p	
-5	35	(1+2)
0	65	
11.5	170	
19.5	275	
27.5	427	

Sulfur dioxide (SO_2) + Strontium iodide (SrI_2)

Ephraim, 1919

t	p	
0	90	(4+1)
6	130	
13	186	
19	250	
25	330	

Sulfur iodide (SO_2) + Barium iodide (BaI_2)

Ephraim, 1919

t	p		t	p	
15	65	(2+1)	0	380	(4+1)
29	145		2.5	438	
41	390		7	560	
49	750		9	623	
			12.5	760	

Sulfur dioxide (SO_2) + Aluminum chloride ($AlCl_3$)

Ephraim, 1919

t	p		t	p
0	20	(1+1)	38.5	105
9.5	40		62	160
20.5	65		100	250

Tesei, 1941

%	f.t.	%	f.t.
0.54	-35	10.24	-15
0.99	-29	15.83	-14
2.65	-18	20.08	-13
7.22	-15	23.51	-11
8.02	-15		

%	λ					
	-10°	-5°	0°	5°	10°	15°
0.54	6.00	6.00	5.88	5.55	4.64	3.78
0.99	5.80	6.00	5.82	5.68	6.09	4.42
2.66	4.50	4.91	5.00	5.07	5.07	4.76
7.79	3.05	2.25	2.42	2.55	2.58	2.63
8.72	2.02	2.27	2.47	2.70	2.82	2.91
11.41	1.62	1.94	2.18	2.38	2.60	2.61
18.81	0.77	0.95	1.15	1.33	1.48	1.61
25.13	0.56	0.75	0.93	1.08	.123	1.33
30.73	0.33	0.45	0.58	0.70	0.79	0.90
	20°	25°	30°	35°	40°	
0.54	2.88	1.50	2.06	1.03	0.66	
0.99	3.58	1.96	2.77	1.45	1.09	
2.66	4.22	3.20	3.71	2.64	2.14	
7.79	2.52	2.28	2.42	2.02	1.74	
8.72	2.89	2.65	2.82	2.48	2.20	
11.41	2.63	2.54	2.59	2.39	2.21	
18.81	1.69	1.72	1.73	1.68	1.60	
25.13	1.45	1.56	1.52	1.55	1.49	
30.73	0.97	1.10	1.04	1.14	1.15	

%	κ					
	-10°	- 5°	0°	5°	10°	15°
0.54	1.56	1.62	1.6	1.5	1.3	1.1
0.99	2.98	2.97	3.0	2.9	2.65	2.3
2.66	5.40	5.9	6.5	6.6	6.6	6.3
7.79	8.00	9.0	9.7	10.2	10.6	10.8
8.72	8.70	10.0	11.1	11.9	12.7	13.1
11.41	9.40	11.3	12.9	14.1	15.0	15.7
18.81	7.40	9.2	11.2	18.1	14.7	16.1
25.13	7.20	9.7	12.1	14.2	16.3	18.1
30.73	5.20	7.3	9.3	11.3	13.1	14.8
	20°	25°	30°	35°	40°	
0.54	0.8	0.6	0.4	0.3	0.2	
0.99	1.9	1.5	1.1	0.8	0.6	
2.66	5.9	5.7	4.5	3.7	3.0	
7.79	10.6	10.2	9.6	8.7	7.5	
8.72	13.3	13.1	12.5	11.7	10.6	
11.41	16.1	16.1	15.8	15.1	14.2	
18.81	17.1	17.7	17.8	17.5	16.9	
25.13	19.7	20.8	21.5	21.6	21.1	
30.73	16.3	17.5	18.7	19.5	20.0	

Sulfur dioxide (SO_2) + Aluminum bromide ($AlBr_3$)

Ephraim, 1919

t	p	
16	11	(1+1)
52.5	122	
75	215	
87.5	292	

Sulfur dioxide (SO_2) + Aluminum iodide (AlI_3)

Ephraim, 1919

t	p	
17.5	10	(1+1)
40	54	
54	96	
78	212	
90	265	

Sulfur dioxide (SO_2) + Germanium tetrachloride ($GeCl_4$)

Bond and Crone, 1934

%	sat.t.	overc. f.t.	%	sat.t.	overc. f.t.
100.00	-	-51.8	41.86	-6.2	-
98.80	-75.0	-53.9	37.78	-7.4	-
97.97	-61.9	-55.0	34.22	-9.0	-
97.07	-53.0	-	31.95	-10.1	-
96.41	-46.6	-	27.52	-13.1	-
93.72	-31.0	-	17.52	-24.6	-
92.72	-27.2	-	13.16	-32.7	-
91.76	-24.2	-	11.31	-36.4	-
90.58	-21.1	-	10.07	-41.2	-
87.42	-15.0	-	7.00	-51.5	-
85.54	-12.3	-	6.26	-55.2	-
82.24	-9.2	-	5.15	-60.5	-58.0
78.62	-7.1	-	4.24	-67.1	-61.2
75.36	-6.1	-	3.48	-73.0	-64.5
69.42	-5.2	-	3.07	-76.1	-66.0
68.78	-5.2	-	2.71	-79.2	-66.8
67.59	-5.0	-	2.17	-86.0	-71.8
63.98	-4.9	-	2.07	-88.0	-72.8
63.06	-4.9	-	1.55	-	-76.5
59.87	-4.7	-	1.48	-	"
58.81	-4.7	-	0.94	-	-76.3
52.37	-4.9	-	0.72	-	-75.9
49.92	-5.1	-	0.00	-	-75.5
45.00	-5.5	-			

Sulfur dioxide (SO_2) + Titanium tetrachloride ($TiCl_4$)

Bond and Stephens, 1929

%	sat.t.	overc.f.t.	%	sat.t.	overc.f.t.
100.00	-	-24.4	43.69	11.0	-
98.86	-	-26.0	37.67	9.6	-
97.47	-	-27.7	33.87	8.3	-
96.37	-	-29.1	30.47	6.5	-
96.31	-42.2	-28.9	26.17	3.5	-
94.95	-33.7	-30.1	23.56	1.5	-
90.65	-13.7	-	21.09	-1.5	-
88.66	-6.3	-	19.92	-3.3	-
85.35	-0.3	-	15.02	-11.2	-
85.16	+0.1	-	14.44	-12.2	-
82.43	4.0	-	12.26	-16.8	-
79.35	6.8	-	9.53	-23.7	-
73.26	10.2	-	8.23	-27.9	-
68.08	11.5	-	7.05	-31.4	-31.4
65.40	11.8	-	6.44	-36.1	-32.7
61.80	11.9	-	4.15	-	-36.1
60.46	12.0 C.S.T.	-	3.41	-	-39.3
58.40	"	-	2.86	-	-46.5
58.05	"	-	2.01	-	-54.2
56.28	11.9	-	1.64	-	-57.8
52.82	11.8	-	0.96	-	-66.8
48.20	11.5	-			

overc.=overcooled

Sulfur dioxide (SO_2) + Titanium tetrabromide ($TiBr_4$)

Bond and Crone, 1934

%	sat.t.	overc.f.t.	%	sat.t.	overc.f.t.
100.00	-	38.2	48.58	98.9	-
99.10	-	35.4	42.28	94.1	-
98.08	-	33.1	36.06	87.6	-
98.07	23.5	30.8	32.46	83.3	-
97.03	34.5	-	29.33	79.7	-
95.15	45.3	-	26.28	74.9	-
95.14	53.4	-	24.22	70.8	-
94.10	57.8	-	21.06	64.7	-
93.20	67.6	-	19.24	61.2	-
91.96	72.0	-	25.92	50.8	-
91.43	81.4	-	12.52	44.0	-
88.73	84.5	-	11.47	40.3	-
88.05	92.6	-	10.66	36.1	-
84.23	97.8	-	9.88	33.8	-
80.82	100.8	-	9.20	31.6	-
78.54	102.9	-	8.14	29.0	29.4
74.90	103.2	-	7.45	22.0	26.2
73.57	103.3	-	7.03	18.8	24.4
71.97	103.5	-	6.36	16.0	23.5
71.08	103.7	-	4.48	-	18.5
70.05	103.8 C.S.T.	-	4.06	-	14.0
68.87	"	-	3.00	-	8.2
65.22	"	-	2.54	-	3.5
64.53	103.3	-	1.39	-	-1.0
63.78	103.0	-	1.22	-	-12.8
60.63	102.4	-	1.14	-	-13.2
57.20	101.4	-	1.08	-	-17.0
53.18	100.8	-	0.80	-	-21.7
52.24	100.4	-	0	-	-75.5
51.61		-			

Sulfur dioxide (SO_2) + Zirconium tetrachloride ($ZrCl_4$)

Bond and Stephens, 1929

mol%	f.t.
1.38	0
1.99	10
2.58	20

Sulfur dioxide (SO_2) + Stannic tetrachloride ($SnCl_4$)

Bond and Beach, 1926

%	f.t.	sat.t. (overcooled)	%	f.t.	sat.t. (overcooled)
100.00	-32.7	-	66.06	-43.8	-45.0
95.23	-38.7	-	61.54	-43.9	-45.1
92.66	-40.6	-	61.35	-43.9	-45.1
86.16	-42.6	-52.0	58.54	-44.0	-45.3
82.41	-43.25	-46.8	55.24	-44.1	-45.9
79.43	-43.6	-45.7	50.76	-44.3	-46.3
79.05	-43.6	-45.7	44.27	-44.8	-48.8
76.07	-43.4	-45.6	31.96	-47.1	-57.5
75.76	-43.4	-45.4	26.29	-49.8	-
72.62	-43.6	-44.9	11.78	-63.0	-
71.75	-43.7	-44.9	5.40	-72.7	-
67.47	-43.8	-45.0 C.S.T.			

Sulfur dioxide (SO_2) + Stannic tetrabromide ($SnBr_4$)

Bond and Beach, 1926

%	f.t.	sat.t.	%	f.t.	sat.t.
100.00	29.45	-	64.00	-	48.50
99.96	28.50	-	63.30	-	48.50
98.74	24.90	-	52.83	-	46.30
98.68	24.80	-	43.60	-	41.67
97.30	21.15	-	43.34	-	41.45
96.24	18.80	-	40.63	-	39.75
95.71	18.23	-	26.33	-	25.29
95.70	18.23	-	26.09	-	25.00
95.62	18.00	-	21.67	16.3	18.42
95.42	17.80	-	20.39	-	16.60
94.12	16.60	-	16.55	13.2	8.00
94.07	16.55	-	14.08	10.7	1.8
93.67	-	18.33	11.88	7.1	-17.8
92.75	-	23.60	10.61	5.45	-
91.65	-	28.70	9.45	3.25	-
89.81	-	34.25	8.93	2.5	-
85.54	-	42.60	6.29	-3.5	-
82.49	-	45.75	5.37	-5.0	-
82.38	-	45.80	2.97	-17.1	-
80.31	-	47.25	2.41	-21.8	-
78.97	-	47.70	1.88	-24.00	-
78.44	-	47.80	0.92	-37.7	-
75.14	-	48.60	0.58	-45.7	-
71.38	-	48.60	0.39	-53.05	-
64.77	-	48.48 C.S.T.			

Sulfur trioxide (SO_3) + Lead oxide (PbO)

Schenk and Rassbach, 1908

%	f.t.	E	tr.t.	
100	882	-	-	
95.2	867	820	-	
92.6	849	"	-	
90.9	837	"	-	
88.9	820	"	-	E
87	842	"	450	
83.3	866	818	"	
80.0	880	-	"	(1+4)
79.1	918	880	"	
71.4	945	"	"	
69	948	941	"	(1+3)
66.7	942	938	"	(1+3) + (1+2)
65.4	951	"	"	
63.7	960	941	"	
62.5	960	939	"	
60.6	965	-	-	(1+2)
58.8	960	951	841	
57.1	-	950	840	
55.6	-	"	839	
54.0	-	952	"	
52.6	-	950	843	
51.9	-	952	"	
51.5	-	953	845	
50.6	-	950	846	
50.0	-	"	845	
	-	-	850	

Jaeger and Germs, 1921

mol %	f.t.	E	tr.t.	
100	879	-	-	
90.9	840	835	-	
82.5	888	835	-	
81	897	835	897	(1+4)+(1+3)
80	953	-	897	
76.7	960	-	-	
75	961	950	-	(1+3)
71.4	950	950	-	(1+3)+(1+2)
69	970	950	-	
66.6	977	-	-	(1+2)
62.5	970	960	864	(1+1) II-I
33.3	960	"	"	
58.8	1006	"	"	
55.5	1092	"	"	
51.3	1157	"	"	
50	1170	-	-	

Lander, 1949

mol%	f.t.	E	mol%	f.t.	E
100	-	878	80.6	928	895
90.9	853	834	80.0	940	"
90.1	860	835	79.4	945	893
88.5	878	"	78.1	954	892
87.0	888	"	77.5	956	888
85.5	895	832	76.3	960	886
84.7	893	830	75.2	-	961
84.0	893	820	73.0	957	950
83.3	899	895	71.9	-	"
82.6	903	"	69.0	972	"
82.0	915	"	33.3	-	975
			(1+3)		

Sulfuryl chloride (SO_2Cl_2) + Stannic tetrachloride ($SnCl_4$)

Locket, 1932

mol%	d	η
	25°	
100	2.2164	814.5
73.2	2.1054	773.9
43.7	1.9580	739.6
20.5	1.8170	714.4
0	1.6573	685.0

Silicon oxide (SiO_2) + Lithium oxide (Li_2O)

van Klooster, 1910

%	f.t.		E
50.8	1215		-
49.9	1243	(1+2)	-
46.6	1212		991
44.9	1159		1015
41.5	-		1010
38.7	1083		1008
35.2	1174		1012
33.3	1188	(1+1)	-
28.9	1157-1135		-
24.9	1083		948
21.2	1009		947
17.9	985		941
14.8	956		937
11.9	-		948
9.2	1081		940

Rieke and Endell, 1910

%	mol%	f.t.	E	
30	46.3	1152	950	
33.3	50.0	1180	-	(1+1)
35.0	52.0	1170	1035	
40.0	57.3	1058	1025	(2+3)
42.5	59.8	998	998	
45.0	62.1	1070	1000	
47.5	64.5	1180	995	
50.0	66.7	1215	-	(1+2)
52.2	68.7	1200	-	

Ballo and Dittler, 1912

mol%	f.t.	E	E	
68.7	1200	-	-	
66.66	1217	1217	-	(1+2)
64.5	1180	995	995	
62.1	1070	1000	1000	
59.8	998	998	998	E
57.3	1058	1028	1028	
52.0	1170	1035	1035	
50.0	1180	1180	-	(1+1)
46.3	1152	950	950	
40.2	1080	"	"	
36.3	1045	955	955	
33.5	955	"	"	
32.1	960	945	945	
30.7	975	930	930	
27.7	1000	920	920	
26.6	1040	945	945	
21.4	-	948	948	
16.9	-	940	940	
0	1600	1600	-	

Kracek, 1930 and 1939

%	f.t.	
49.1	1255	Li_2O + (1+2) + L
44.7	1024 E	(1+2) + (1+1)
33.22	1201	(1+1)
19.9	1033	(1+1) + (2+1) + L
17.8	1028 E	(2+1) + Tridymite
9.0	1470	Tridymite + Cristobalite
0.0	1713	Cristobalite

%	n glass	%	n glass
44.68	1.567	19.92	1.535
33.22	.557	19.01	.533
25.12	.546	17.71	.529
23.36	.543	15.02	.522
22.07	.540	10.03	.503
21.00	.538	0.00	.459

%	f.t.	E	
54.0	1255	-	
47.93	1226	1024	(1+2) + (1+1)
46.30	1143	1024	
44.68	1028	-	
42.10	1095	1024	
39.50	1145	1024	
34.63	1199	-	
33.22	1201	-	(1+1)
30.87	1195	-	
30.00	1188	-	
25.12	1125	-	
23.36	1092	-	
22.07	1070	-	
21.00	1052	1033	
19.92	1034	1033	
19.07	1032	-	
17.71	1037	-	(2+1) + tridymite
15.02	1217	1028	
12.56	1340	1028	
10.03	1434	1028	
7.52	1535	-	Cristobalite
5.04	1581	-	
3.0	1622	-	

Kracek, 1939 (fig.)

%	f.t.	E	tr.t.
43.4	1092	1033	-
43.8	1068	"	-
44.1	1055	"	-
44.5	1034	"	936
44.8	1031-1032	1028	-
45.1	1039	"	-
46.0	-	"	-
46.6	-	"	-

Bockris, Tomlinson and White, 1956

mol%	d	$\tau \cdot 10^5$	mol%	d	$\tau \cdot 10^5$
		1400°			
20	2.169	3.4	45	2.074	10.4
25	.153	5.5	50	.055	10.7
30	.138	7.2	55	.017	11.5
35	.122	7.7	60	1.980	12.6
40	.098	9.0	65	.955	13.2

Austin, 1947

X-ray diffraction (see author)

Silicon oxide (SiO_2) + Sodium oxide (Na_2O)

Morey and Bowen, 1925

mol%	f.t.	
49.8	1086.5	(1+1)
45.1	1031.0	
44.2	1001	
38.8	863	
-	835	E
37.5	847	
36.9	859	
35.2	871	
33.33	873.5	(2+1)
33.30	873.0	
32.6	872.5	
32.1	868	
28.6	831	
26.7	832	Quartz
25.0	830	"
24.2	841	
19.0	1145	Tridymite
11.1	1457	
-	1470	tr.t. Cristobalite-Tridymite
4.9	1596	Cristobalite
0	1710	

D'Ans and Löffler, 1930

mol%	f.t.	m.t.	E	tr.t.
66.3	1075	-	-	960
65.1	1057	-	-	"
64.9	1048	-	-	"
63.9	1000	989	-	958
62.4	1062	-	-	959
61.3	1090	-	-	958
60.4	1114	1114	-	-
59.3	1110	1110	-	-
55.8	1069	-	1016	-
53.2	1065	1061	1015	-
52.0	1074	1064	1014	-
50.0	1088	1086	-	-

mol%	f.t.	
100	850	
-	800	E (1+2) + Na_2O
66.7	1083	(1+2)α
66.7	960	tr.t. (1+2)α - (1+2)β
63.5	1000	E (1+2)α + (2+3)
60	1122	(2+3)
54.5	1015	E (2+3) + (1+1)
50	1088	(1+1)
38.5	840	E (1+1) + (2+1)
33.3	874	(2+1)
25.9	793	E (2+1) + SiO_2 (Quartz)
0	1710	Cristobalite

Kracek, 1939 (fig.)

%	f.t.	E	tr.t. 1	tr.t. 2	
29.6	955	-	-	-	
28.6	900	-	-	-	
27.8	860	837	707	678	
27.3	848	"	"	"	
25.1	874	-	"	"	
24.1	-	-	"	"	
23.7	-	-	"	"	
21.3	830	-	-	-	(1+1)
21.3	805	-	707	678	
20.1	790	-	-	-	
20	825	789	-	-	
18.2	905	769	-	-	

Foex, 1944

%	f.t.	tr.t. 1	2	3	4	5
0	1713	1470	870	768	678	-
10	1470	"	"	"	"	-
20	870	-	"	"	"	-
21	793	-	-	"	"	E SiO_2 +(2+1)
35	874	-	-	-	"	706 (2+1)
41	846	-	-	-	"	" E(2+1)+(1+1)
52	1089	-	-	-	-	" (1+1)
59	1022	-	-	-	-	- E(1+1)+Na_2O
63	1118	-	-	-	-	-

Bockris, Tomlinson and White, 1956

mol%	d	$\tau \cdot 10^5$	mol%	d	$\tau \cdot 10^5$
		1400°			
11	2.216	0.3	47.5	2.196	12.5
15	.221	3.5	47.6	.194	12.7
20	.222	4.4	48.7	.193	13.4
24.6	.218	6.2	50	.192	12.8
30	.210	8.7	52.5	.191	13.1
35	.202	10.3	52.5	.185	14.0
40	.201	10.7	55	.177	13.7
45	.195	11.9	55	.180	13.7
45	.197	11.5	55	.179	13.7
46.2	.199	11.9	57.5	.178	13.8
47.5	.200	12.8	60	.163	14.3

Tillotson, 1918

vol%	n_D	vol%	n_D
0	1.464	29.30	1.5110
14.50	.4865	36.20	.5137
19.45	.4950	45.80	.5200
22.00	.5000		

Silicon oxide (SiO_2) + Potassium oxide (K_2O)

Kracek, Bowen and Morey, 1929

%	f.t.	E	tr.t.	
0	1713	-	-	cristobalite
4.3	1635	-	-	
19.8	-	815	-	
-	-	-	1470	
15.0	1254	-	-	tridymite
16.8	1149	-	764	(5+1)
18.4	1038	-	"	
19.5	935	-	"	
20.5	905	-	765	
20.8	-	834	"	
-	-	-	870	
20.8	840	-	-	Quartz
21.9	765	-	-	(4+1)
22.1	764	-	-	
22.5	762	-	-	
23.0	760	-	-	
23.7	753	749	-	(2+1)
24.3	-	745	-	
24.0	794	-	-	
24.2	825	-	-	
24.5	854	-	-	
24.9	893	-	-	
25.1	910	-	-	
25.4	923	-	-	
25.8	933	-	-	
26.0	954	-	-	
26.7	974	-	-	
26.9	980	-	-	
27.7	988	-	-	
28.1	992	-	-	
28.4	1005	-	-	
29.1	1009	-	-	
29.6	1029	-	-	
30.2	1031	-	-	
30.5	1036	-	-	
30.7	1035	-	-	
31.2	1034	-	-	
31.8	1027	-	-	
32.3	1016	-	-	
32.6	1000	-	-	
32.9	984	-	-	
33.1	968	-	-	
33.5	940	-	-	
33.5	943	-	-	
33.7	910	-	-	
33.8	895	-	-	
33.9	882	-	-	
34.3	835	-	-	
34.3	819	-	-	
34.6	805	-	-	
35.1	783	-	-	

Morey, 1917

mol %	f.t.
50.0	976 (1+1)
43.5	775 E
33.3	641 (2+1)
20.0	525 E
-	575 tr.t.

Bokris, Tomlinson and White, 1956

mol%	d	$\tau.10^5$	mol%	d	$\tau.10^5$
		1400°			
10	2.216	0.0	30	2.156	11.6
15	.201	3.4	35	.150	12.8
20	.184	7.0	40	.136	14.0
25	.167	9.6	45	.113	16.0

Joffè, 1952

Dielectric losses (see author) .

Silicon oxide (SiO_2) + Cuprous oxide (Cu_2O)

Otin, 1912

%	mol %	f.t.	E	tr.t.	d
100	100	1208	1088	-	5.744
96.58	92.23	1111	1092	-	5.566
95.75	90.47	1163	1053	942	5.596
94.42	87.69	1090	1062	-	5.272
93.33	82.61	1094	1042	-	4.863
89.43	78.08	1165	1033	-	5.508
88.26	75.99	1082	1046	923	5.309
84.92	70.36	1200	1048	-	4.995

Berezhnoi, Karyakine and Dudovski, 1952 (fig.)

%	f.t.	%	f.t.
100	1230	60	1680
90	1065 E	20	1690
80	1150	0	1750

Silicon oxide (SiO_2) + Magnesium oxide (MgO)

Anderson and Bowen, 1914

%	f.t.	
0	1625	
30	1543	(1+1)+SiO_2(Cristobalite)
35	1543	
36	1543	
36	1552	SiO_2(Cristobalite)
38.8	1556	
38.8	1557	(1+2)+(1+1)
39	1557	
39.2	1557	
40.1	1557	
42	1557	
42	1625	(1+2)+L
50	1557	(1+2)+(1+1)+L
57.2	1890	(1+2)
65	1850	E(1+2)+MgO
100	2800	

Bowen and Anderson, 1914

mol%	f.t.	
100	2800	
68.0	1850	(1+2) + MgO
66.7	1890	(1+2)
49.4	1557	(1+2)+(1+1)
46.6	1543	(1+1)+SiO_2
0	1625	

Greig, 1927 (fig.)

%	f.t.		E	tr.t.
0	1710		-	-
1	1700		1545	-
2	1700	glass 1 + glass 2	"	-
15	1697		"	-
	1694	Cristobalite + glass 2	"	-
26	1694		"	-
	1710	glass 1 + glass 2	"	-
30	1700		"	-
31	1692	Cristobalite + glass 2	"	-
	1696	glass 1	"	-
32	1667	glass	"	-
	1654		"	-
35	1545	E	"	-
39	1560	tr.t. (1+1) + (1+2)	"	1560
40	1560	(1+1)	-	"
60	1890	(1+2)	-	"

Wartenberg and Prophet, 1932 (fig.)

%	f.t.	E
0	1630	-
20	1625	-
30	1600	-
35	1550	1550
40	1560	(1+1)
45	1750	-
56	1880	(1+2)
60	1850	-
65	1950	-
100	2800	-

Silicon oxide (SiO_2) + Calcium oxide (CaO)

Jänecke, 1912

%	f.t.		E	tr.t. 1	tr.t. 2
0	1600		-	1190	-
10	1580		1418	"	-
20	1540		"	"	-
30	1480		"	"	-
37	1418		"	"	-
40	1453		"	"	-
45	1500		"	"	-
48.2	1512	(1+1)	-	-	-
50	1508		1433	1190	680
52	1484		"	"	"
54	1433	E	"	"	"
55.5	1503		"	1405	"
65	2082	(1+2)	"	"	"
67.5	2015	E	2015	"	"
75	2380		"	"	"

Rankin, 1915

%	f.t.	tr.t. 1	tr.t. 2	
0	1625	-	-	
48.3	1540	-	-	(1+1)
37	1436	-	-	E(1+1) + SiO_2(Tridymite)
58.2	1475	-	-	(2+3)
54.5	1455	-	-	E(2+3) + (1+1)
65.0	2130	1420	675	(1+2)
73.6	1900	-	-	(1+3)
100	2870	-	-	

Greig, 1927 (fig.)

%	f.t.	E	tr.t.	
0	1710	-	-	Cristobalite
2	1700	1430	1470	
5	1699	"	"	
10	1700	"	"	
15	"	"	"	
20	"	"	"	
26.5	"	"	"	
28	"	"	"	
31.8	1650	"	"	
35	1480	"	"	
35.5	1470	"	"	Cristobalite
37	1430	"	-	Tridymite
39.5	1470	"	-	
48	1540	-	-	
55	1460	1460	-	(1+1+ + (2+3)
56	1475	"	1475	(2+3) + (1+2)
58.5	1475	-	1475	(2+3)

Bowden, Schairer and Posnak, 1933 (fig.)

%	f.t.	E	tr.t.	
0	1710	-	-	
2	1700	1430	1470	
10	"	"	"	
20	"	"	"	
28	"	"	"	
34	1550	"	"	
35	1470	"	"	
36	1430	"	-	
48	1540	-	-	(1+1)
50	1530	1450	-	
55	1450	"	-	E
56	1475	"	1475	tr.t.(1+1) + (2+3)
58	1475	-	"	(2+3)
65	-	-	-	(1+2)

Kolobova, 1941

%	f.t.	%	f.t.
56.5	1625	46.27	1554
54.5	1460 E	44.0	1545
53.0	1507	42.0	1525
52.0	1527	37.0	1440 E
51.0	1542	35.0	1470
50.27	1550	33.0	1580
50.0	1550	31.0	1640
48.27	1555	29.0	1680
48	1556		
		(1+1)	

%	f.t.	%	f.t.
90.8	2450	67.21	2050
80.5	2320	66.25	2085
73.85	2220	65.18	2110
70.20	2125	63.65	2075
69.15	2100	58.32	1800

Bockris and Lowe, 1954

wt%	mol%	f.t.		E
29	30.5	1700		1470
33	34.6	1610		1470
37.2	38.8	1436		1436
37.2	38.8	1436		"
37.2	38.8	1436		"
40	41.6	1470		"
42	43.7	1490		"
47	48.7	1535		1455
48	49.7	1540	(1+1)	"
49	50.7	1540		"
50	51.7	1520		"
51	52.7	1510		"
52	53.7	1500		"
53	54.7	1490		"
54.5	56.1	1455		"
56	57.6	1550		1475

wt%	mol%	η 1450°	1500°	1550°	1600°
29	30.5	-	-	-	-
33	34.6	-	-	-	-
37.2	38.8	2110	1440	1020	730
37.2	38.8	2070	1400	980	710
37.2	38.8	2120	1400	970	710
40	41.6	-	935	648	468
42	43.7	-	765	560	405
47	48.7	-	435	317	241
48	49.7	-	-	288	218
49	50.7	-	-	270	200
50	51.7	-	-	243	181
51	52.7	-	303	220	166
52	53.7	-	288	211	157
53	54.7	-	257	139	140
54.5	56.1	3311	239	180	139
56	57.6		-	-	113

wt%	mol%	η 1650°	1700°	1750°	1800°
29	30.5	-	1360	1040	850
33	34.6	1000	780	605	450
37.2	38.8	525	392	310	-
37.2	38.8	513	393	300	250
37.2	38.8	525	390	315	-
40	41.6	357	275	216	180
42	43.7	295	235	195	150
47	48.7	190	150	120	99
48	49.7	168	133	106	88
49	50.7	152	119	97	80
50	51.7	139	111	90	75
51	52.7	128	101	83	72
52	53.7	120	96	79	66
53	54.7	110	90	75	66
54.5	56.1	105	81	68	60
56	57.6	90	74	62	54

Silicon oxide (SiO_2) + Diopside ($CaMgO_6Si_2$)

Bowen, 1915

%	f.t.	
100	1396	
90	1360	
-	1365	glass + Diopside
85	1360	
83	-	glass + Tridymite
80	1448	glass + Cristobalite
75	1577	
0	1625	

Silicon oxide (SiO_2) + Anorthite ($CaAl_2O_8Si_2$)

Andersen, 1915 (fig.)

%	f.t.	
44	1470	tr.t. (Cristobalite, tridymite)
52	1352	E
60	1397	
70	1445	
80	1490	
90	1525	
100	1550	

Silicon oxide (SiO_2) + Strontium oxide (SrO)

Eskola, 1922

%	f.t.	
40	1636	glass
"	1600	glass + Cristobalite
53.82	1368	glass + Tridymite
"	1361	
"	1356	
50	1363	(1+1) + glass
50	1430	glass + (1+1)
"	1420	
60	1571	
"	1552	
63.22	1584	glass + (1+1)
"	1580	
"	1575	(1+1)
65	1554	glass + (1+1)
"	1546	
"	1538	E
67	1544	(1+1) + (2+1)
"	1617	glass + (2+1)
"	1600	
72	1628	
77.46	1634	(2+1)
100	1750	

%	d	n F	Tl	D	C
46.2	3.201	1.591	1.587	1.584	1.581
50	-	1.598	1.595	1.591	1.589
60	-	1.632	1.627	1.624	1.621
63.2	3.537	1.640	1.636	1.632	1.629
67	-	1.652	1.648	1.644	1.641

Greig, 1927 (fig.)

%	f.t.	tr.t.	E	
0	1710	-	-	Cristobalite
2	1700	1470	1350	
15	1694	"	"	
30	1692	"	"	
38	1611	"	"	
43.5	1470	"	"	tr.t.Cristobalite Tridymite
46	1350	-	1350	E SiO_2 + (1+1)
55	1500	-	"	(1+1)
63	1580	-	"	"

Silicon oxide (SiO_2) + Barium oxide (BaO)

Eskola, 1922 (fig.)

%	f.t.
0	1710
10	1705
20	1690
30	1645
40	1555
43	1475 tr.t. (cristob.- tridym)
47	1374 E tridum + (2+1)
57	1420 (2+1)
60	1430-1440 mixcrystals (3+2) + (2+1)
62.9	1450 (3+2)
65	1437 E (3+2) + (1+1)
72	1605 (1+1)
74.5	1551 E (1+1) + (1+2)
78	1640

spectral lines	n					
	25.0%	26.9%	29.2%	29.8%	32.6%	33.3%
α (F)	1.602	1.602	1.629	1.622	1.625	1.627
α (Tl)	.599	.599	.615	.619	.622	.623
α (D)	.597	.597	.612	.616	.619	.620
α (C)	.595	.595	.610	.613	.616	.617
β (F)	1.617	1.618	1.621	1.627	1.631	1.632
β (Tl)	.614	.614	.618	.623	.627	.628
β (D)	.612	.612	.615	.620	.624	.625
β (C)	.610	.610	.612	.617	.621	.622
γ (F)	1.632	1.644	1.645	1.647	1.652	1.652
γ (Tl)	.625	.639	.641	.642	.647	.648
γ (D)	.621	.636	.638	.639	.644	.645
γ (C)	.618	.633	.635	.636	.640	.641

Greig, 1927 (fig.)

%	f.t.	E	tr.t.	
0	1710	-	-	Cristobalite
5	1688	1370	1470	
10	1683	"	"	
15	1679	"	"	
25	1674	"	"	
35	1636	"	"	
40	1583	"	"	
45	1489	"	"	
45.5	1470	"	"	tr.t. Cristobalite Tridymite
46.5	1370	1370	-	E SiO_2 + (2+1)
56	1430	"	-	(2+1) and tr.t.
63	1450	-	-	(3+2)
65	1440	1440	-	
72	1600	"	-	

Austin, 1947

X-ray diffraction (see author) .

Silicon oxide (SiO_2) + Zinc oxide (ZnO)

Bunting, 1930

mol%	f.t.	mol%	f.t.
0	1695	60	1500
10	1700 L_1+L_2	70	1512 (1+2)
20	" "	77	1507 tr.t.
30	" "	80	1560
36	" "	90	1760
40	1600 E	100	1975
50	1435		

Silicon oxide (SiO_2) + Lead Monoxide (PbO)

Cooper-Shaw-Loonis, 1909

mol%	f.t.	E
100	888	-
90	853	720-727
87	825	722-729
80	723	723
77.78	727	718
75.00	716	716
71.43	732	718
69.00	742	714
66.67	746	- (1+2)
65.00	742	717
60.00	722	718
57.14	723	716
52.00	764	718
50.00	766	- (1+1)
48.00	765	-
42.86	760	-

Weiller, 1911

%	f.t.	E	%	f.t.	E
100	876	-	89	-	615
96	777	675	88	-	645
95	736	676	87	-	637
93.5	705	675	86	-	616
92	-	675	85	-	600
90	715	657			

Cooper, Kraw and Klein, 1912

%	f.t.	%	f.t.
72.90	760	93.70	753
78.80	768 (1+1)	95.47	780
84.25	716 E	97.10	850
88.11	759 (1+2)	100	889
91.75	715 E		

Geller, Creamer and Bunting, 1934 (fig.)

%	f.t.	
100	886	PbO
97.5	842	"
95	775	"
93.3	725	"
-	720	(1+4)α + (1+4)β
91.8	714	E
90	730	
88.0	743	(1+2)
84.6	716	E
82.5	745	
79	764	(1+1)
77.5	763	
75	760	
72.5	755	
70.4	732	E
70.0	860	Quartz
69.0	872	Quartz and Tridymite
67.5	910	Tridymite

Mc Murdie and Bunting, 1939

(1+1) , (1+2) , (1+4)

Weiller, 1911

%	d	%	d
91.8	8.06	75.91	5.74
91.4	8.124	72.44	5.28
87.0	7.005	71.00	5.20
84.0	6.94	67.66	4.87
82.2	6.45	66.22	4.73
79.92	6.14		

%	n			
	Hg red	Hg orange	D	Hg yellow
87	2.058	-	2.078	2.083
84	-	2.025	2.028	2.038
82.2	-	-	1.979	1.986
79.72	-	-	1.960	-
75.91	-	1.922	1.930	-
72.44	-	-	1.864	1.866
29.00	-	-	1.770	1.779

%	n		
	Hg green	Hg blue	Hg violet
87	2.095	2.174	-
84	2.018	2.104	2.112
82.2	-	2.048	-
75.91	1.935	1.965	-
72.44	-	1.897	-
29.00	1.785	1.811	1.824

Shartis, Spinner and Smock, 1948

%	σ			
	700°	800°	900°	1000°
96.80	-	-	134.4	142.0
99.83	-	-	147.8	157.7
91.64	-	-	173.6	176.5
86.82	-	-	183.7	186.5
84.70	-	187.3	192.4	194.6
83.24	-	-	-	199.4
82.50	193.2	196.8	199.1	202.2
75.70	-	-	217.1	204.4
72.88	-	-	224.5	225.6
71.69	-	-	-	227.8
69.92	-	-	-	231.3
69.84	-	-	-	230.0
67.26	-	-	-	233.7
65.09	-	-	-	233.0

%	σ			
	1100°	1200°	1300°	1400°
96.80	145.4	154.2	158.0	161.3
99.83	152.1	158.0	164.1	162.7
91.64	179.2	182.0	183.3	-
86.82	189.6	193.9	197.3	-
84.70	196.1	-	-	-
83.24	202.2	205.7	210.2	-
82.50	204.4	208.2	209.4	-
75.70	221.0	222.6	223.4	-
72.88	225.7	226.5	227.2	-
71.69	228.4	227.5	228.5	-
69.92	232.0	232.7	230.5	229.1
69.84	230.8	230.8	231.7	232.2
67.26	235.6	235.6	234.5	235.1
65.09	234.0	234.0	235.0	235.7

Joffé, 1952 (fig.)

mol%	ε
	at room temperature
35	13
40	15
50	18.5
55	20
60	20.1

Silicon oxide (SiO_2) + Manganous oxide (MnO)

Doerinckel, 1911

mol%	wt%	f.t.	E
80	82.2	1321-1280	-
75	77.9	1320-1275	-
60	73.2	1320-1280	-
66.7	70.2	1327-1280	-
63.4	66.7	1318	1170
60	63.2	1260	1185
58.35	62.2	1240	1180
55	59	-	1190
52.5	56.5	1220	1185
50	54.1	1230	-
47.5	51.6	1240-1210	-
46	50	1240-1220	-
45	49	1265-1210	-
40	43.9	1290-1220	-

Cain, 1920

%	f.t.	%	f.t.
5	above 1500	50	1268
10	" "	55	1264
15	" "	60	1256
20	" "	65	1228
25	" "	70	1258
30	" "	75	1287
35	" "	80	1282
40	" "	85	1284
45	" "	90	above 1500
		95	" "

White, Howat and Hay, 1933

wt %	mol %	f.t.
100	100	1785
63.4	66.7	1330 (1+2) Tephroite
62.5	58.5	1208 E (1+1) + (1+2)
54.1	50.0	1280 (1+1) Rhodonite
50 - 2	-	above 1630 $L_1 + L_2$
0	-	1705

%	f.t.	m.t.
90.8	-	1350
72.5	1440	1240
68.0	1360	1160
53.0	1525	1210

Greig, 1927

10% sat.t.= 1725°

Silicon oxide (SiO_2) + Ferrous oxide (FeO)

Keil and Dammann, 1925 (fig.)

%	f.t.	%	f.t.
100	1395	43	1000 E
96	1075 E	37	1260
71	1503 (1+2)	27	1200
60	1115 E	0	1460
50	1280 (1+1)		quartz

Greig, 1927 (fig.)

%	f.t.	%	f.t.
0	1710	30	1690
3	1690	40	1690
10	1690	42	1650
20	1690		

Bowen and Schairer, 1932

%	f.t.	
100	1380	FeO
90.83	1315	
79.45	1215	
76	1177	E
74.52	1193	
72.46	1202	
70.39	1205	(1+2) : Fayalite
68.35	1203	
63.90	1187	
62	1178	E
61.41	1260	SiO_2 (Tridymite)
60.16	1365	
57.59	1475	Cristobalite
56.40	1515	

Bowen, Schairer and Posnyak, 1933 (fig.)

%	f.t.	tr.t.	E
0	1710	-	-
2	1690	1470	1180
20	"	"	"
40	"	"	"
50	1640	"	"
57	1500	"	"
58	1470 tr.t.	"	"
60	1370	-	"
61	1260	-	"
61.2	1180 E	-	"
71	1205 (1+2)	-	"
77	1175 E	-	1175
80	1220	-	"
90	1310	-	"
100	1380	-	"

Silicon oxide (SiO_2) + Nickel monoxide (NiO)

Greig, 1927

10% sat.t. 1725° (sic.)

Silicon oxide (SiO_2) + Cobaltous oxide (CoO)

Greig, 1927

10% sat.t.= 1725° (sic.)

Silicon oxide (SiO_2) + Aluminum oxide (Al_2O_3)

Neumann, 1910

%	f.t.	%	f.t.
100	2000	29.7	1710
92.9	1960	25.3	1690
82.0	1920	22.0	1670
81.9	1888	17.5	1650
62.9	1850	14.5	1630
56.0	1825	12.8	1620
50.5	1790	10.1	1600
45.8	1770	8.4	1670
40.4	1750	6.5	1710
36.1	1730	0	1770

Rankin, 1915

%	f.t.	
100	2050	
-	1816	(1+1)
64	1810	E (1+1) + Al_2O_3
13	1610	E (1+1) + SiO_2
0	1625	cristobalite

Bowen and Grieg, 1924 (fig.)

%	f.t.	
0	1700	Cristobalite
5.5	1545	E
10	1665	
15	1705	
25	1740	
40	1750	
50	1800	
55	1810	tr.t. (2+3) + Corundum
72	1810	(2+3)
72	1910	Corundum
80	1950	
90	2000	
100	2050	

Sosman, 1933

%	f.t.	
0	1728	
5.5	1549	E
55	1830	tr.t.
100	2040	

tr.t: 1470° tridymite-cristobalite

Toropov and Galakhov, 1951 (fig.)

%	f.t.		%	f.t.	
0	1710		50	1840	
5	1540	E	60	1850	
10	1650		72	1870	(2+1)
20	1715		78	1850	E
30	1740		90	2000	
40	1790		100	2050	

Silicon oxide (SiO_2) + Titanium dioxide (TiO_2)

Bunting, 1933

%	f.t.	%	f.t.
0	1715	40	1780
5	1665	50	1790
10	1575	60	1800
11	1540	70	1810
12	1625	80	1815
15	1680	90	1820
20	1720	100	1825
30	1760		

Silicon oxide (SiO_2) + Zirconium dioxide (ZrO_2)

Zirnowa, 1934

mol%	f.t.	mol%	f.t.	
100	2715	32.00	2060	
95.32	2680	30.08	2020	
95.00	2670	28.00	2000	
84.11	2570	25.00	1935	
78.07	2510	23.85	1925	
76.80	2485	19.74	1885	
70.00	2405	8.87	1805	
67.54	2360	6.96	1795	
61.14	2275	6.09	1785	
54.55	2330	4.00	1745	(1+1)
50.00	2430	3.00	1705	
47.85	2420	0.72	1710	
40.31	2340	0	1715	

Toropov and Galakhov, 1956

%	f.t.	t.quenching	
75	2400	-	
67	2350	2450	glass
		2380	ZrO_2 + glass
62	2280	2400	glass
		2320	ZrO_2 + glass
57	2280	2350	glass
		2300	glass + glass 2
52	-	2450	glass
		2370	glass + glass 2
		2280	glass + glass 2
		2240	ZrO_2 + glass
47	-	2460	glass
		2420	glass + glass 2
		2400	glass + glass 2
		2320	glass + glass 2
		2200	ZrO_2 + glass
43	2230	2400	glass + glass 2
		2350	glass + glass 2
		2260	ZrO_2 + glass
40	2260	2350	glass
		2300	glass + glass 2
		2200	ZrO_2 + glass
35	2200	2300	glass
		2230	ZrO_2 + glass
30	2180	-	-

$L_1 + L_2$ 47 % 2430° C.S.T.

Arsenic trisulfide (As_2S_3) + Thallous sulfide (Tl_2S)

Canneri and Fernandes , 1925

mol %	f.t.	E
100	448.5	-
90	427	295
80	418	"
70	376.5	"
60	341	"
50	-	"
45	291	"
40	278	-
37	-	200
35	251	"
32.5	317	-
30	303.5	215
28	-	"
25	300	-
23	269.5	-

(2+1)

Arsenic trisulfide (As_2S_3) + Silver sulfide (Ag_2S)

Jaeger and van Klooster, 1912

%	f.t.	m.t.	$tr.t._1$	E	$tr.t._2$
100	-	842	-	-	176
95	-	715	454	469	-
90.1	-	610	463	"	-
85.1	-	502	469	"	-
83.1	469	469	-	-	-
80.1	-	481	467	469	-
75.1	-	490	-	-	-
70.1	-	479	323	399	-
65.2	-	445	339	399	-
60.2	399	398	-	-	-
55.2	-	414	394	399	-
50.2	417	417	-	"	-
40.2	413	413	-	"	-
33.5	410	410	-	"	-
25.1	408	407	-	"	-
20.1	407	404	-	"	-
10.1	397	351	-	"	-

Silicon disulfide (SiS_2) + Silver sulfide (Ag_2S)

Cambi, 1912

%	f.t.	E	min.	
100	835	-	-	
99.32	817	798	50	
98.90	804	799	100	
97.82	848	799	70	
95.01	933	798	40	
92.90	944	798	10	(1+4)
91.49	958	798	-	
91.24	958.9	-	-	
90.75	958	-	-	(2+3)
90.01	957	746	20	
89.05	955	745	30	
86.83	904	743	-	
84.76	808	747	70	
82.78	765	749	80	
81.26	751	748	120	
80.60	755	749	90	
80.02	756	-	-	
79.96	754	-	-	
78.96	742	678	10	
77.34	728	675	30	
76.67	714	678	90	
75.45	681	680	130	
74.43	692	681	100	
72.90	698	678	100	
69.91	711	676	80	

Silicon disulfide (SiS_2) + Lead sulfide (PbS)

Cambi, 1912

%	f.t.	tr.t.1	min	tr.t.2	min	tr.t.3	min
100	1114	-	-	-	-	-	-
95	-	-	-	745	10	-	-
91.3	-	766	10	747	10	-	-
89.1	887	764	15	748	30	-	-
88.5	880	761	20	748	50	-	-
86.8	855	765	15	747	20	-	-
85.0	828	765	30	751	20	-	-
83.9	817	766	80	-	-	-	-
83.0	804	758	40	-	-	739	10-5
80.6	785	758	30	-	-	737	10
80.5	774	765	10	-	-	736	40
79.5	756	-	-	-	-	737	50
78.7	748	-	-	-	-	737	30
76.5	704	-	-	-	-	-	-

(1+3) (1+2) (2+3)

Arsenic triselenide (As_2Se_3) + Mercuric selenide (HgSe)

Pélabon, 1908

mol%	f.t.	
100	880	
95	790	(2+3)
66	365	

Nitric acid (HNO_3) + Potassium nitrate (NO_3K)

Groschuff, 1904

%		f.t.	%		f.t.
24.4		-6	47.2		+22.5
32.6		+14	47.8		23.5
34.8		17	48.6		25.5
37.2		19.5	49.4		27
44.5	(2+1)	22	50.1		29
47.8		21.5	50.9	(1+1)	30.5
48.6		21.5	49.4		21
50.9		20	50.9	NO_3K	39
37.2	E	-4	51.7		50
44.5		+16.5			

Pacault and Chédin, 1950

%	χ	%	χ
0	0.316	38.5	0.322
9.0	.315	41.4	.322
18.3	.315	100	.329
26.8	.318		

Sulfuric acid (H_2SO_4) + Lithium sulfate (Li_2SO_4)

Kendall and Landon, 1920

mol%	f.t.		mol%	f.t.
0	+10.3		22.11	105.2
0.81	8.0		23.45	109.3
2.85	2.0		23.92	114.7
4.63	-3.6		24.73	117.7
6.42	+3.0		26.41	125.4
7.63	6.5		28.95	135.8
8.75	9.7		31.28	142.5
9.73	11.4		32.30	144.5
11.24	12.9		33.31	147.1
12.32	13.1		33.97	148.6
12.50	13.3	(7+1)	35.12	152.4
12.85	13.0		36.01	155.2
13.02	13.0		38.33	159.4
13.04	17.9	(2+1)	40.11	161.3
14.91	32.6		40.76	161.5
15.31	36.7		42.60	164.7
15.87	40.9		45.10	167.8
16.23	44.3		46.00	168.5
16.58	48.9		47.60	186.5
17.01	55.3		48.10	229.9
18.39	70.9	(1+1)	48.55	256.0
19.32	82.7		48.79	258.6
20.06	91.2		49.85	279.9
21.16	99.0		50.48	289.0
			51.80	315.0

Sulfuric acid (H_2SO_4) + Lithium hydrogen sulfate ($LiHSO_4$)

Gillespie and Wasif, 1953

m	d	η
	25°	
0.2491	1.8310	28070
.5712	.8455	32740
.6331	.8484	33780
1.127	.8660	42820

Sulfuric acid (H_2SO_4) + Sodium sulfate (Na_2SO_4)

Cambi and Bozza, 1923

wt%	mol%	f.t.	tr.t. I	tr.t. II	tr.t. III
8.48	6.01	29.7 (9+2)	-	-	-
12.43	8.93	44.6	-	-	-
13.89	10.02	48.9	-	-	-
17.15	12.49	53.6	-	-	-
18.50	13.55	54.2	-	-	-
19.24	14.12	55.7	-	-	-
20.03	14.74	60.7	54.4(2+1)	53.0(1+1)	-
21.68	16.04	71.9	54.6	61.7	-
22.82	17	-	-	67.6	-
24.41	18.20	84.0	54.4	74.4	-
26.50	19.93	92.0	"	-	-
28.00	21.16	96.0	53.6	102.6(1+1)	-
32.30	24.04	104.6	53.0	103.5	-
33.35	25.67	116.0	-	101.5	-
36.06	28.03	129.2	-	98.2	-
37.59	29.37	133.0	-	101.6	-
39.04	30.65	135.8	-	-	-
39.42	31.0	138.0	-	98.0	90.5
42.01	33.33	148.5	-	99.0	-
43.78	34.95	155.0	-	-	-
49.73	40.58	169.0	-	-	-
52.22	43.00	172.0	-	-	-
59.15	50.0	185.7	-	-	-
61.06	51.98	178.0	-	-	177.0
61.56	52.50	207	-	-	"
63.82	54.90	231	-	-	177.4

Kendall and Landon, 1920

mol%	f.t. I	f.t. II	mol%	f.t.
0	10.4	-	29.38	139.8
1.98	4.3	-	30.43	148.2
3.69	10.8	-	31.37	148.2
4.79	21.2	-	33.99	159.0
5.84	28.9	-	34.28	160.6
7.04	36.8	-	37.27	170.0
8.91	45.4	-	41.27	177.8
10.38	49.9	(9+2)	43.41	180.7
12.63	55.0	-	44.12	180.9
15.11	57.4	-	48.25	185.5
16.29	71.5	62.9	50.23	185.8
17.48	78.6	70.4	50.97	190.9
19.21	88.5	79.7	51.37	194.3
20.95	95.7	87.0	53.96	223.3
23.14	102.1	93.6	54.60	230.2
25.36	112.5	-	55.74	243.2
26.37	122.3	-	56.72	254.2
28.90	136.5	-	58.83	274.5
(9+2)	(2+1)	(1+1)	(x+1)	

Pascal and Ero, 1923

%	f.t.	%	f.t.	m.t.
15.00	+14	53.25	176	-
18.04	36	56.20	180	-
21.54	56	59.16	185	179.5
24.70	75	59.98	180.5	174
28.84	97	61.20	176.5	"
27.50	97	62.22	176	"
35.42	120	63.24	178.5	174.5
41.41	145.5	64.06	245	"
47.33	166	65.29	250	173.5
	(3+1)	(2+1)	(1+1)	

Sulfuric acid (H_2SO_4) + Sodium hydrogen sulfate ($NaHSO_4$)

Gillespie and Wasif, 1953

m	d	η
	25°	
0.2755	1.8437	27490
.4392	.8531	29450
.6968	.8665	33700
.9108	.8784	35910
1.365	.8994	44190
1.454	.9064	46130

Sulfuric acid (H_2SO_4) + Potassium sulfate (K_2SO_4)

Kendall and Landon, 1920

mol%	f.t.	mol%	f.t.
0	+10.4	35.97	162.0
0.30	9.5	37.56	171.5
1.10	7.2	38.28	176.4
2.38	+3.4	39.22	178.3
4.23	-3.2	40.85	188.4
5.51	-7.9	41.73	195.3
6.87	-0.8	43.31	200.2
8.34	+15.4	46.77	210.0
10.40	36.5	47.94	213.8
11.98	48.9	48.62	215.7 (1+1)
13.53	59.1	50.06	218.6
15.39	69.6	50.36	218.2
17.21	77.1	51.70	214.1
19.03	83.3	52.19	213.1
21.80	89.2	53.14	212.6
24.05	91.5 (3+1)	53.27	212.5
25.53	91.6	53.58	219.8
28.27	89.0	54.52	235.9
29.29	105.4	54.82	235.6
30.93	122.1	55.24	237.5
31.37	126.7	55.71	264.3
32.49	136.7	56.59	296.8
33.86	149.5	58.47	300

Cambi and Bozza, 1923

wt%	mol%	f.t.	E	min	tr.t.	
					1	2
0	-	+10.4	-	-	-	-
6.43	3.72	-1.3	-	-	-	-
9.95	6.01	-9.8	-	-	-	+0.1
11.71	6.95	+0.1	-	-	-	+0.1
16.07	9.73	30.8	-	-	-	-
18.73	11.48	44.2	-	-	-	-
22.43	13.45	60.6	-	-	-	-
25.22	16.05	71.0	-	-	-	-
27.66	17.71	78.5	-	-	-	-
32.02	20.95	87.6	-	-	-	-
36.91	24.76	91.6	-	-	-	-
37.41	25.16	91.6 (3+1)	-	-	-	-
39.76	27.09	90.2	88.65	15	-	-
41.77	28.75	92.0	88.35	36	-	-
45.19	31.69	124.0	-	-	-	-
48.20	34.41	146.1	88.8	27	-	-
49.57	35.62	155.8	89.2	21	-	-
50.70	36.66	161.5	88.6	-	-	-
52.06	38.16	167.1	89.15	20	149	-
54.81	40.57	181.6	-	-	153	-
57.47	43.10	193.2	-	-	153	182
59.95	45.73	201.1	-	-	142	185
60.33	46.12	203.0	-	-	-	-
62.81	48.71	211.0	-	-	-	-
63.99	50	213.8 (1+1)	-	-	-	-
65.40	51.55	208.8	205	6	157	-
66.15	52.49	205.9	205	11	157	-
66.68	52.96	208.0	205.1	11	159.9	-
67.11	53.45	218.1	205.1	9	158.5	-
68.00	54.47	234.0	-	-	155.1	-
68.89	55.48	255.0	205	-	158.0	-
69.82	56.56	290.5	-	-	-	-

Sulfuric acid (H_2SO_4) + Potassium hydrogen sulfate ($KHSO_4$)

Gillespie and Wasif, 1953

m	d	η
	25°	
0.2526	1.8446	25230
.4633	.8578	26140
.5291	.8622	26890
.8084	.8780	28300
1.1990	.8993	31930
1.7760	.9276	38840

Sulfuric acid (H_2SO_4) + Silver sulfate (Ag_2SO_4)

Kendall and Davidson, 1920

mol%	wt%	f.t.	
		I	II
0	0	10.4	-
1.73	0.55	7.0	-
1.99	0.64	6.1	-
3.67	1.19	2.4	-
5.69	1.86	-1.5	-
7.41	2.46	+13.0 (2+1)	-
9.03	3.03	24.5	-
10.17	3.44	29.9	-
11.19	3.81	33.0	-
11.80	4.04	35.3	-
12.38	4.26	37.6	-
11.66	3.99	-	25.3 (1+1)
11.80	4.04	-	28.5
12.38	4.26	40.2	34.1
12.88	4.44	43.1	38.6
13.63	4.73	50.4	-
13.88	4.82	-	46.1
14.57	5.09	-	50.4
14.64	5.12	-	51.7
15.62	5.50	59.3	-
16.76	5.96	-	62.3
17.43	6.23	65.3	65.3
18.71	6.75	-	70.1
21.00	7.72	-	78.6
23.38	8.76	-	85.1
23.96	9.02	-	86.7
26.13	10.01	-	92.0
28.55	11.11	-	96.0
30.94	12.35	-	101.3
33.74	13.81	-	105.1
35.80	14.92	-	108.0
38.57	16.49	-	112.8
42.45	18.83	-	117.1
44.69	20.27	-	120.0
46.70	21.61	-	122.2
47.17	21.93	-	122.8
47.22	21.97	-	138.9
47.75	22.32	-	144.6
48.07	22.56	-	165.3
48.49	22.81	-	171.8
48.69	22.99	-	187.2

Sulfuric acid (H_2SO_4) + Calcium sulfate ($CaSO_4$)

Kendall and Davidson, 1920

mol%	f.t. I	f.t. II	mol%	f.t. I	f.t. II
0	10.4	-	5.31	27.8	22.2
0.98	9.1	-	5.74	-	28.4
1.99	7.2	-	5.92	36.9	31.5
3.17	4.7	-	6.63	45.4	39.7
4.00	2.2	-	7.11	51.3	45.4
4.45	-	3.4	7.59	55.1	49.4
4.70	-	10.1	8.34	60.3	-
4.84	19.1	13.5	8.66	63.2	-
5.19	-	18.4			

Sulfuric acid (H_2SO_4) + Barium sulfate ($BaSO_4$)

Kendall and Davidson, 1920

mol%	f.t.	mol%	f.t.	
0	10.4	6.70	+5.7	(3+1)
0.95	9.1	7.05	9.3	
2.12	6.6	7.30	12.4	
3.50	2.8	8.09	20.0	
5.46	-3.2	8.28	21.5	
6.50	-6.6	8.59	23.0	
6.70	-7.3			

Phosphoric acid (PO_4H_3) + Sodium diorthophosphate (NaH_2PO_4)

Parravano and Mieli, 1908

% (1+1)	f.t.
52.72	98.5
69.59	111.0
77.55	119.0
81.71	122.0
87.20	123.0

Phosphoric acid (PO_4H_3) + Potassium diorthophosphate (KH_2PO_4)

Parravano and Mieli, 1908

% (1+1)	f.t.
18.17	38.5
58.42	84.0
77.53	110.0
89.26	126.5

TO THE MEMORY OF MY FATHER AND OF MY MOTHER

Notice for Users

1. Scope of the work

The data compiled refer only to binary systems, concentrated solutions.

As components, I have accepted all kinds of substances, elements or compounds, with the exception of metallic alloys, a category covered by many other books.

As concentrated solutions, I choose to consider arbitrarily systems between 10 and 90 per cent by weight; I left also out of consideration data relating to dilute solutions, if there is only one measure between 10 and 20 % .

All data, so far as possible, have been reproduced from the original publications, if available; in other cases, the actual source of the data is given in the bibliographic reference. Preference has been given to the experimental data, rather than to values interpolated from a formula; in many cases we had to read the data from graphs, with help of a grating (this is denoted by "fig").

2. General Plan

All data are classified by systems, since values of different properties may help to caracterise their physical nature.

The systems have been arranged in four categories, one for each volume of this book, as follows:

A. Both components are organic compounds, excepting the hydroxyl derivatives.

B. Both components are organic compounds, one at least being a hydroxyl derivative.

C. One at least of the components is a metallic compound.

D. All other systems.

In that volume are also included the general table of bibliographic references and the general table by substances.

I consider as non-metals the following twenty elements:

B - C,Si - N,P,As - O,S,Se,Te -
H,F,Cl,Br,I - He,Ne,Ar,Kr,Xe

I call non-metallic compounds those with only these elements; and organic compounds all such compounds with at least one atom of C. As metallic compounds, I consider all those with at least one metallic atom. Ex.: CSi is an organic compound, sodium benzoate a metallic one, and HCl a non-metallic one.

3. Order of the systems

In each section, the binary mixtures are brought together in great divisions, according to the degree of physico-chemical similitude of their components; for ex., in the third volume, the first part deals with mixtures of two metallic salts, the second one with solutions of metallic salts in water and the third, with solutions of these salts in all other solvants, non-metallic or organic.

In each of these divisions, the binary mixtures are listed, according to the order of the first component, and, for each of them, according to the order of the second component; for ex., all systems with methane come first, methane + butane being listed before methane + benzene, since butane comes before benzene in my classification.

a) For organic compounds, the general order is: hydrocarbons, halogen derivatives, oxygen derivatives (excluding the hydroxyl ones), nitrogen, mixed oxygen and nitrogen derivatives, and last the hydroxyl derivatives of any kind.

In each of these groups, the aliphatic derivatives come first (saturated and then unsaturated), then the polymethylenes, the aromatic compounds and finally the heterocyclic ones.

The sulfur derivatives are listed after the corresponding oxygen ones, the phosphorus, after the nitrogen ones, the silicon and boron after the carbon ones. In each group, the derivatives produced by halogen substitution are placed at the end of the respective group; for ex., ethylenchlorhydrin comes at the end of the alcohol group.

In accordance with this rule, we have the following arrangement:

Hydrocarbons: paraffins, ethylenic and acetylenic hydrocarbons, polymethylenes and aromatic hydrocarbons.

Halogen derivatives: derivatives of the same hydrocarbon are grouped together, in order of the number of hydrogen atoms substituted by halogen atoms, fluorine derivatives first, then chlorine, bromine and iodine derivatives.

Oxygen derivatives: first the ether oxides, with open chain (ethyl ether) or closed ring (dioxane), the aldehydes and ketones, the anhydrides, and finally the esters.

Nitrogen derivatives: nitriles and amines.

Mixed Oxygen and Nitrogen derivatives: compounds of the amide type, and then nitroso- and nitro- derivatives.

Hydroxyl derivatives: first the alcohols and oximes, then the phenols and finally the acids.

N.B. The presence in the molecule of a chemical function listed later, relegates this compounds to the end of that category, for ex., acetoacetic esters come after the esters.

b) Metallic Compounds. Most of them are electronic compounds which are classified as follows:

The salts, oxides, sulfides, etc. come together, so long as the metal has the same electrovalency, for ex., the ferrous compounds are classified with nickel, cobalt, manganese ones. but the ferric compounds, with aluminum and chromic salts.

The metallic ions are classified in series of the same electrovalency, according to the periodical table:

Li, Na, K, Rb, Cs, Tl^{+} - Cu^{+}, Ag, Au^{+}, Hg^{+}

Be, Mg, Ca, Ba, Sr, Sn^{++}, Pb^{++} - Zn^{++}, Cd^{++}, Hg^{++}, Cu^{++}, Mn^{++}, Fe^{++}, Ni^{++}, Co^{++}

Al, Ga, In, Tl^{+++}, Cr^{+++}, Fe^{+++}, Rare Earths - Sb^{+++}, Bi^{+++}.

Ge, Ti, Th, Sn^{++++}, Pb^{++++} - Uranyl.

For each metallic ion, the salts are arranged according to the valency of the anion and the oxygenated salts after all others, as follows:

fluorides, chlorides, bromides, iodides, cyanides, thiocyanates, etc.;

oxides, sulfides, selenides, etc. - nitrides, borides, carbides, silicides;

hydrates, thiohydrates - nitrites, chlorites...

chlorates, bromates, iodates, nitrates;

phosphites, arsenites;

perchlorates - permanganates;

phosphates, arsenates, etc.;

carbonates, sulfites, metasilicates;

sulfates, selenates, chromates, manganates;

orthosilicates.

4. Order of the constants.

So far as possible, especially for systems where the data are particularly numerous, the order in which the properties are classified is as follows:

a) Heterogeneous equilibria:

Critical constants; saturates vapour pressure for the triphase equilibrium.

Vapour pressure curve; boiling curve and azeotropes.

Composition of liquid and vapour coexisting phases.

Densities of coexisting phases and rectilinear diameter.

Composition of the two liquid phases and eventually of the saturated vapour; critical solution point.

Freezing and melting curve; eutectic and transition points.

Equilibria of the condensed phases under high pressure.

b) Properties of phases: first for the gas, then the liquid and finally the mixed crystals:

Densities, coefficients of expansion and of compressibility.

Viscosity and surface tension.

Refractive index and optical dispersion.

Dielectric constant; electrical conductivity.

Optical rotatory power.

Magnetic rotation; magnetic susceptibility.

c) Thermal constants:

Specific heat; heat of solution or mixing.

Heat of vaporization and fusion.

Thermal conductivity.

5. Choice of units.

So far as possible, we have always used units of the c.g.s. system; when necessary, we have converted the original results into these units, so far as it did not involve the use of a coefficient whose value has changed sometimes. Ex: we could, without any ambiguity, transform specific volumes into densities, or density d_t^t into d_4^t; but to transform molar concentration in weight concentration, if not made by the author himself, would have involved a somehow arbitrary choice of atomic weights.

All our numerical data have been taken as given in the original paper; we always gave priority to direct experimental results, rather than recalculated curves.

Here follow some additional details about the choice of units:

Viscosity: in poises . 10^5

Surface tension: in dynes/cm

Temperature: t in centigrade; T = absolute temperature = t + 273.16

Pressure: p - in mm Hg; P - in atmospheres; P_{kg} - in kg/cm^2

π and τ represent pressure or temperature coefficient of the constant considered, which means its change by kg or by degree; but when it relates to volume changes, π and τ are coefficients of compressibility or expansion, as given by the formulae:

$$v_t = v_o \cdot (1 + \tau . t) \quad \text{and}$$

$$v_P = v_1 \cdot (1 - \pi . P)$$

Specific heat: in calories / gram of mixture

Heat of mixing, heat of vaporization, etc. - in calories / mole of mixture.

In case other units were exceptionally used, this is expressely stated in column headings.

N.B. Scientists of the whole world always agree to give their results in units of the metric system; only in Anglo-Saxon countries, did some authors give also their results in British units, for the ease of their technicians. But in recent years some American physico-chemists, namely Sage and his co-workers, have published in Industrial and Engineering Chemistry some extensive tables of data on isotherms of mixtures of hydrocarbons, only in British units (°F, pressure in Lb/sq.in., etc.), without any corresponding tables in metric values, which makes them quite unsuitable for general use in other countries. We have made in most cases the necessary calculations to reproduce these data in metric units, but this work is so laborious and tedious that we were unable to give the complete data; and we wish to protest here with energy against this new mode of publication, which takes no notice of the international scientific public.

6. Nomenclature and bibliographical data.

A. Nomenclature.

Inside this work the common names of the substances are used, with their molecular formulae; but in the Table at the end of the 4th volume, they are classified in the same order as in the Chemical Abstracts, with the different synonyms. For ex., the compound we call ethylene chloride in the book itself, is also named: 1,2-dichlorethane, in the table.

B. Bibliographical data.

Inside the book, the data are reproduced under the name of their author, with the year of publication. The complete bibliographical reference is to be found in the alphabetical list of authors, at the end of this book.

For the transcription of Russian names, we have applied the rules used in Chemical Abstracts. But in case of a Russian author, all of whose quoted publications have been printed in Latin caracters, we have reproduced his name as he had it transcribed himself; when necessary, we give also in the list of authors, the alternative transcription of his name.

7. Symbols and abbreviations.

α	Rotatory power, for the length = 10 cm
(α)	Specific rotatory power
$(\alpha)^{mol}$	Molar " "
$(\alpha)_{magn}$	Specific magnetic rotatory power
$(\alpha)^{mol}_{magn}$	Molar " "
ε	Dielectric constant
η	Viscosity, in poises $(.10^5)$*
$\varkappa$	Specific conductivity $(.10^4)$
λ	Equivalent conductivity
π	Pressure coefficient $(.10^6)$
σ	Surface tension, in dynes/cm
τ	Temperature coefficient
χ	Magnetic susceptibility $(.10^6)$ (specific)
C	Crystal
C.S.T.	Critical solution temperature
C.V.T.	" vaporization "
D	Diffusion coefficient $(.10^5)$
D_{therm}	Thermal diffusion coefficient
D b.t.	Boiling temperature difference
D f.t.	Freezing " "
Dp	Pressure difference
Dt	Temperature "
Dv	Volume "
E	Eutectic
L	Liquid
M	Molarity
N	Normal concentration
P	Pressure, in atmospheres
P_{kg}	" in kg/cm^2
Q comb	Heat of combustion (cal/gram mixture)
Q dil	" dilution (cal/mole mixture)
Q diss	" dissolution "
Q melt	" fusion "
Q mix	" mixing "
Q trans	" transition "
Q vap	" vaporization "
R	Resistivity
S	Solid
T	Absolute temperature
U	Specific heat (cal/gram mixture)
V	Vapour
aq	Aqua, water
atm	Atmosphere
b.t.	Boiling temperature
c	g/100 cc solution
cc	Cubic centimeter
cal	Calorie (small)
crit.	Critical
d	Density (t/4)
dissoc.	Dissociation
e	Electromotive force (in volts)
f.t.	Freezing temperature
g	Gram
l	Liter
m	Molality
mm	Millimeter
mg	Milligram
min	Minutes
mol	Molar
m.t.	Melting temperature
n	Refractive index
p	Pressure in mm Hg
sat.t.	Saturation temperature (mutual solubility)
sol	Solution
s. or sym.	Symmetrical
t	Temperature, centigrade
tr.t.	Transition temperature
trans.	Transition
vol	Volume
v_o	Volume at 0%
w.l.	Wave length (in Ångström unit)
%	Weight percent
I, II, etc.	Polymorphic forms
I - II	Transition of form I into form II

* The given powers for some units are systematically used in the Tables, unless otherwise stated.

SYMBOLS AND ABBREVIATIONS

α	Rotatory power, for the length 10 cm
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Q dil	" dilution (cal/mole mixture)
Q diss	" dissolution "
Q melt	" fusion "
Q mix	" mixing
Q trans	" transition "
Q vap	" vaporization "
R	Resistivity
S	Solid

T	Absolute temperature
U	Specific heat (cal/gram mixture)
V	Vapour
aq	Aqua, water
atm	atmosphere
b.t.	Boiling temperature
c	g/100 cc solution
cc	Cubic centimeter
cal	Calorie (small)
crit.	Critical
d	Density (t/4)
dissoc.	Dissociation
e	Electromotive force (in volts)
f.t.	Freezing temperature
g	Gram
l	Liter
m	Molality
mm	Millimeter
mg	Milligram
min	Minutes
mol	Molar
m.t.	Melting temperature
n	Refractive index
p	Pressure in mm Hg
sat. t.	Saturation temperature (mutual solubility)
sol	Solution
s. or sym.	Symmetrical
t	Temperature, centigrade
tr. t.	Transition temperature
trans.	Transition
vol	Volume
v_0	Volume at 0%
w.l.	Wave length (in Ångström unit)
%	Weight percent
I,II,etc.	Polymorphic forms
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